


Cite this: *RSC Adv.*, 2020, 10, 44408

Received 24th October 2020
Accepted 20th November 2020

DOI: 10.1039/d0ra09088c

rsc.li/rsc-advances

A click-based modular approach to introduction of peroxides onto molecules and nanostructures†

Alissa Horn and Patrick H. Dussault*

Copper-promoted azide/alkyne cycloadditions (CuAAC) are explored as a tool for modular introduction of peroxides onto molecules and nanomaterials. Dialkyl peroxide-substituted alkynes undergo Cu(I)-promoted reaction with azides in either organic or biphasic media to furnish peroxide-substituted 1,2,3-triazoles. Heterolytic fragmentation of the peroxide to an aldehyde, a side reaction that appears to be related to the formation of the triazole, can be suppressed by use of excess alkyne, the presence of triethylsilane, or by use of iodoalkyne substrates. Complementary reactions of simple alkynes with azido-substituted peroxides are much less efficient. Click reactions of alkynyl peroxyacetals are also reported; reductive fragmentation can be minimized by increasing the distance between the peroxyacetal and the alkyne. The strategy enables modular introduction of dialkyl peroxides and peroxyacetals onto gold nanoparticles, the first such process to be reported.

Introduction

Organic peroxides are capable of a wide range of reaction pathways,¹ and have been applied as oxidants,² radical initiators,³ pharmacophores,⁴ enzyme inhibitors,⁵ and synthons for electrophilic transfer of alkoxide.⁶ Peroxides are also of interest as intermediates and products of oxidative degradation and sources of reactive oxygen species.⁷ This rich chemistry contrasts with the limited number of reports describing reactivity or application of peroxides at surfaces.⁸ This discontinuity cannot be completely attributed to concerns for stability or safety. Peroxides are compatible with a surprising variety of synthetic transformations,⁹ and many peroxides may be safely employed by following established precautions.¹⁰ A more significant barrier to broader use of peroxides is the lack of effective methods for their modular introduction under mild conditions.

“Click chemistry”, a family of strategies based upon rapid and specific pairwise reactions of matched functional groups, has become an indispensable tool for modular approaches to molecules, supramolecules, and functionalized surfaces and nanoparticles.¹¹ The most widely applied of these transformations, is the copper-assisted azide/alkyne cycloaddition (CuAAC), now a workhorse reaction for ligation and modification of molecules and supramolecules.¹² CuAAC chemistry has been employed to generate derivatives of the peroxide antimalarial artemisinin,¹³ and we became interested in the potential for modular introduction of peroxides *via* click reactions. We

now describe explorations of the scope of CuAAC chemistry reaction in terms of the nature of the peroxide group, whether the peroxide is linked to the alkyne or azide partner, the distance of the peroxide from the reaction center, and the influence of reaction conditions. The fragmentation of the peroxides observed in some of the click reactions of peroxyacetals demonstrates a potential for controlled generation of reactive oxygen species. We also demonstrate successful application of these ligations to modular introduction of peroxides on the surface of functionalized Au nanoparticles (AuNP),

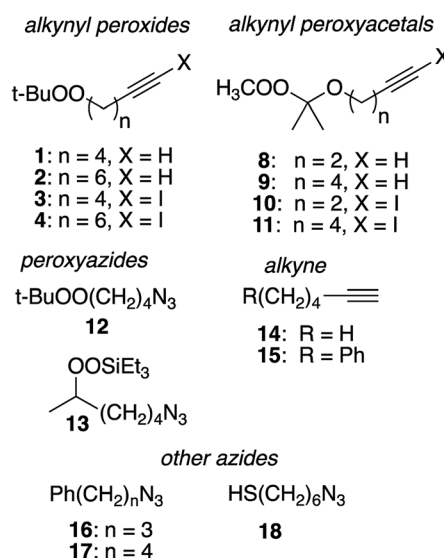


Fig. 1 Substrates used for CuAAC chemistry studies.

Department of Chemistry, University of Nebraska-Lincoln, Lincoln, NE 68588-0304, USA. E-mail: pdussault1@unl.edu

† Electronic supplementary information (ESI) available: ¹H and ¹³C NMR spectra of molecules, as well as ¹H, IR, and XPS characterization of functionalized Au nanoparticles (100 pages). See DOI: 10.1039/d0ra09088c



$$\text{Poly(alkynyl methacrylate-OH)} \xrightarrow[\text{-78 } ^\circ\text{C}]{\text{O}_3/\text{O}_2, \text{CH}_2\text{Cl}_2} \text{Poly(alkynyl methacrylate-OOH)}$$

5: $n = 1$ (3 %) 5a: $n = 1$ (30 %)
 6: $n = 2$ (35 %) 6a: $n = 2$ (10 %)
 7: $n = 4$ (71 %) -

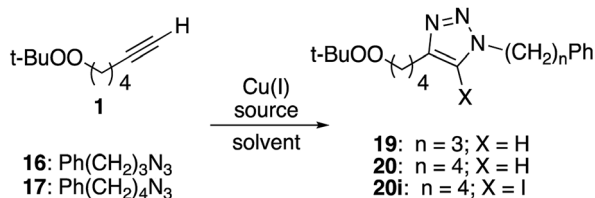
$$\text{Poly(alkynyl methacrylate-OOH)} \xrightarrow{\text{KOt-Bu, MeI}} \text{Poly(alkynyl methacrylate-I)}$$

8: $n = 2$ (66 %) 10: $n = 2$ (81 %)
 9: $n = 4$ (90 %) 11: $n = 4$ (98 %)

Cl(CH2)4Br
 $\xrightarrow{\begin{array}{l} 1. \text{NaN}_3 \text{ (95\%)} \\ 2. \text{NaI, } \Delta \text{ (91\%)} \\ 3. \text{t-BuOOH, CsOH (62\%)} \end{array}}$
CC(C)(C)OC(=O)CCCN=[N+]=[N-]
12

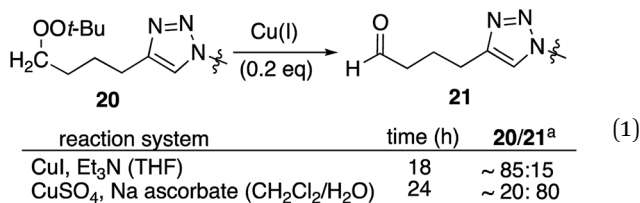
HO(CH2)4CH=CH2
 $\xrightarrow{\begin{array}{l} \text{a. Ph}_3\text{P, I}_2, \text{imid. (75\%)} \\ \text{b. NaN}_3 \text{ (98\%)} \\ \text{c. Co(acac)}_2, \text{O}_2, \text{Et}_3\text{SiH (46\%)} \end{array}}$
CC(C)C(C)CCCN=[N+]=[N-]
13

Table 1 Conditions for click chemistry with peroxy alkynes



Azide	Cu(I) source	Solv.	<i>t</i> (h)	Conv.	Prod. (yield, %)
16	CuCl, NEt ₃	THF	3	89%	19 (25%)
17	CuI, NEt ₃	THF	3	80%	20 (37%) 20i (5%)
17	CuSO ₄ , Na ascorbate	CH ₂ Cl ₂ /H ₂ O	1.5	85%	20 (42%)

Preparation of peroxide-substituted azides is summarized in Scheme 3. Selective nucleophilic substitution of 1-bromo-4-chlorobutane afforded 1-chloro-4-azidobutane,²² which underwent CsOH-promoted substitution with *tert*-butyl hydroperoxide to generate azidoalkyl peroxide **12**. The synthesis of a secondary peroxide/azide began with conversion of 5-hexene-2-ol to the corresponding iodide. Nucleophilic displacement with sodium azide was followed by Mukaiyama peroxidation to generate peroxy azide **13**.²³ As **12** and **13** both incorporate two energy-rich groups within a low molecular weight framework,²⁴ initial syntheses were conducted on modest scales (~1 mmol). Although both molecules

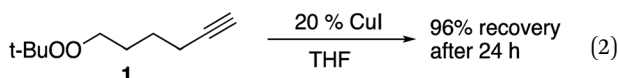


^a Ratio assessed by ¹H NMR.

proved to be stable to at least 100 °C (open chamber DSC, range limited by volatility), subsequent preparations were limited to ≤1 g and we avoided exposing either molecule to elevated temperatures.¹⁰ Preparations of 5-hexynyl benzene (**15**), azidoalkylbenzenes **16** and **17**, and 6-azido hexanethiol (**18**) are detailed in the Experimental section.

Copper(I) catalyzed click reaction with peroxy alkyne

Reaction of peroxyalkyne **1** with azidoalkylbenzenes **16** or **17** under either homogeneous (CuI and NEt₃, in THF) or biphasic (CuSO₄, sodium ascorbate, methylene chloride/water) conditions led to the rapid (≤15 min) appearance of triazoles **19** or **20** (Table 1);^{12,25} the presence of a peroxide was evident based upon a redox-sensitive TLC indicator.²⁶ These initial reactions were stopped after brief reaction periods (1.5–3 h); longer reaction times led to consumption of starting material but also contamination with a nearly inseparable byproduct later established to be the tetrazole aldehyde (**21**, *vide infra*). A small amount of the corresponding 5-iodotriazole **20i**, easily distinguished from **20** by the ¹³C signal for C₅ (120.7 ppm for C–H vs. 78.0 ppm for C–I) was isolated from the reaction employing CuI.²⁷



The relationship between reaction time and aldehyde formation was examined using the reaction of **1** and **17** (Table

2). Ratios of peroxide products (**20**, **20i**) vs. aldehyde (**21**) were established by NMR; samples of pure aldehyde were also isolated for characterization. The near absence of aldehyde for reactions conducted in the presence of excess alkyne (entries 6 and 9) would emerge as an important tool (*vide infra*).

Resubjecting isolated triazole **20** to either set of reaction conditions led to formation of aldehyde **21** (eqn (1)); in the presence of CuSO₄/ascorbate, the aldehyde became the major product.

In an effort to better understand the factors leading to aldehyde formation, we probed the reactivity of triazole **20** towards reaction components (Table 3). No decomposition was observed in the presence of base or CuSO₄ or ascorbate. However, treatment with CuI, CuI/triethylamine, or CuSO₄/Na ascorbate generated significant amounts of aldehyde; the same was true for reaction with either Cu(I) source and azide. No aldehyde was observed in the presence of mixtures containing added alkyne (entries 5, 6, 11 and 12) which did not result in aldehyde formation. Interestingly, the 5-iodotriazole (**20i**) was unaffected by conditions which rapidly degraded **20** (entry 13).

The results in Table 3 suggest that the formation of aldehyde may be accelerated in the presence of the neighboring triazole. Supporting this hypothesis, alkynyl peroxide **1** proved inert towards reaction with copper iodide in THF (eqn (2)).

Consistent with literature reports,²⁸ click reactions proceeded more rapidly in the presence of tris[(1-benzyl-1*H*-1,2,3-triazol-4-yl)methyl]amine (TBTA), a Cu(I) ligand (Table 4). However, aldehyde formation continuing to be a problem and the overall yields of clicked peroxide were no greater, and sometimes less, than for reactions in the absence of the ligand. In contrast, performing the CuI or CuSO₄/ascorbate-catalyzed reactions in the presence of excess alkyne completely suppressed aldehyde formation. This led to the hypothesis that the lack of byproduct might reflect the ability of free alkyne to protonate the C₅-cuprated triazole intermediate in CuAAC chemistry.^{12,29} We also investigated this reaction in the presence of a silane, a functionality also known to quench organocopper species,³⁰ and again observed suppression of aldehyde formation. Finally, replacement of the terminal alkyne with an iodide furnished only the 5-iodoalkyl peroxide; no aldehyde was observed. These reactions, although allowed to proceed for 18–24 h, were typically complete (TLC) within 30 minutes.

Iodoalkynes display enhanced reactivity relative to terminal alkynes in CuAAC reactions,²⁷ and this is also observed in the reactions of the peroxide-substrates. A competition between

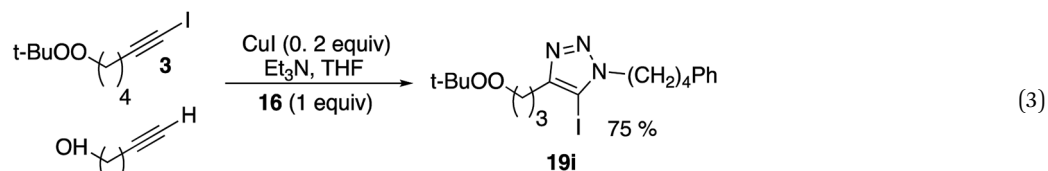
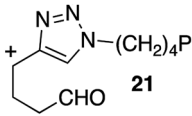


Table 2 Influence of reaction time on byproduct formation

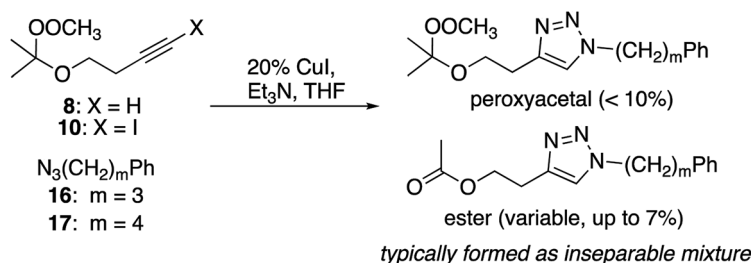
$1 + 17 \xrightarrow[\text{solvent}]{0.2 \text{ equiv Cu(I)}} [20, 20i] + \text{21}$ 				
Entry	Catalyst	Solvent	<i>t</i> (h)	21 (yield ^a , %)
1	CuI/Et ₃ N	THF	1	6%
2	"	"	3	8%
3	"	"	8	13%
4	"	"	18	20%
5	"	"	27	30%
6 ^b	"	"	3	<2%
7	CuSO ₄ /sodium ascorbate	CH ₂ Cl ₂ /H ₂ O	4	20%
8 ^c	"	"	4	38%
9	"	"	18	57%
10 ^b	"	"	1.5	5%
11 ^b	"	"	4	22%

^a Ratios assessed by ¹H NMR. ^b 1.2 equiv. 1. ^c 0.4 equiv. Cu(I).

structurally analogous alkyne derivatives led to preferential consumption of the iodoalkynyl peroxide 3.

Proximity of peroxide and alkyne

Click reactions of peroxyoctyne 2 and the corresponding iodoalkyne (4) produced a good yield of peroxy triazoles 22/23



and iodotetrazoles 22i/23i (Table 5). For reactions employing CuI as the copper source, formation of the aldehyde byproducts 24/25 was minimal (2) or did not occur at all (4). In the case of reactions using CuSO₄, aldehyde formation was significant for reactions involving alkyne 2 and minimal for iodoalkyne 4.

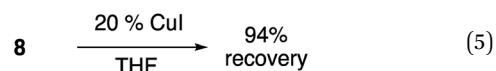
Alkynyl peroxyacetals

Cu(I)-promoted reaction of the butynyl peroxyacetal (8) with azides 16 or 17 gave poor results (eqn (4)). In some cases, we observed successful click reaction to form inseparable mixtures of the desired peroxyacetal and the acetate ester derived from peroxyacetal fragmentation. No improvement was observed based upon the presence of excess alkyne, TBTA, or triethylsilane. In the presence of promoters derived from

Cu(II), the reactions underwent a color change from dark green to light blue after 30–60 minutes. The color change could be temporarily reversed by the addition of more ascorbate but little or no clicked product was detected. Additionally, minimal progress was observed with the corresponding iodoalkynyl peroxyacetal (10).

In contrast, the longer chain peroxyacetal (9) and the corresponding iodoalkyne (11) both reacted with azides in the presence of Cu(I) to give peroxy triazoles as the only major product (Table 6). Once again, no reaction was observed in the presence of a promoter derived from CuSO₄/ascorbate.

A separate control experiment (eqn (5)) made obvious the stability of the alkynyl peroxyacetals towards Cu(I) in the absence of other reaction components.



CuAAC reactions of peroxyalkyl azides

Reactions of simple alkynes with peroxide-containing azide 12 proceeded in modest yield using CuI/Et₃N (Table 7); the



Table 3 Stability of peroxide tetrazoles to reaction components

20: Y = H
20i: Y = I

21

Entry	S. mat	Reagents	Azide (equiv.)	Alkyne (equiv.)	21 (yield ^a , %)
1	20	Et ₃ N	—	—	Nr
2	20	CuI	—	—	95%
3	20	CuI, Et ₃ N	—	—	70%
4	20	"	1.0	—	50%
5	20	"	1.0	1.0	0%
6	20	"	—	1.0	0%
7	20	Na ascorbate	—	—	0%
8	20	CuSO ₄ , Na ascorbate	—	—	43%
9	20	CuSO ₄	—	—	0%
10	20	CuSO ₄ , Na ascorbate	1.0	—	61%
11	20	CuSO ₄ , Na ascorbate	1.0	1.0	0%
12	20	CuSO ₄ , Na ascorbate	—	1.0	0%
13	20i	CuI, Et ₃ N	—	—	0%

^a NMR yields; isolated yield in parentheses.

Table 4 Influence of additives

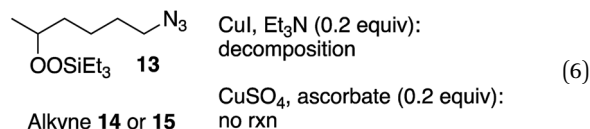
A. CuI/Et₃N, THF;
B. CuSO₄/ascorbate, CH₂Cl₂/H₂O;
C. CuI/Et₃N, DMSO/H₂O
D. CuSO₄/ascorbate, DMSO/H₂O
(all 0.2 equiv Cu)

n = 3: Y = H (**19**)
Y = I (**19i**)
n = 4: Y = H (**20**)

Alkyne	Azide	Cond.	Additive (equiv.)	t (h)	Product (yield, %)
1	17	A	Alkyne 1 (1.2)	18	20 (55%)
1	17	B	Alkyne 1 (1.2)	18	20 (68%)
1	17	C	TBTA (1%)	3	20 (20%)
1	17	D	TBTA (1%)	3	20 (35%)
1	16	A	Et ₃ SiH (1.2)	24	19 (62%)
3	16	A	—	24	19i (65%)

corresponding reactions with CuSO₄/ascorbate provided lower yields and required substantially longer reaction times. In both series, significant amounts of the peroxyazide were recovered. Although analysis of these reactions was complicated by the presence of inseparable byproducts, we could verify that little or no aldehyde was generated under these conditions.

Secondary silyl peroxide **13** failed to give clicked products (eqn (6)); instead, producing a mixture of reduction and fragmentation products.^{1a,31} In the absence of alkyne, **13** was stable to both the CuI or CuSO₄ reagent systems (not shown).

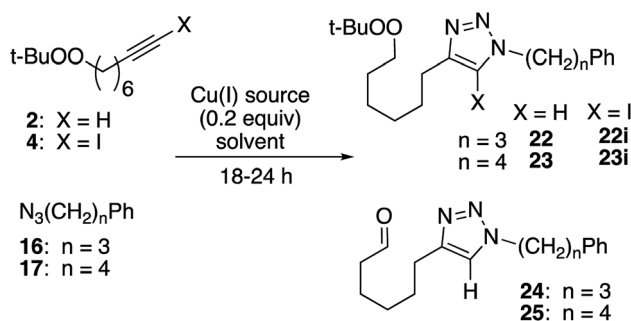


Influence of peroxides on the CuAAC reaction

The lower reactivity and yields observed with the peroxy-substituted azides led us to investigate whether the presence of a peroxide was exerting a general dampening influence on the CuAAC reaction. The reaction of 1-hexyne (**14**) and phenylalkyl



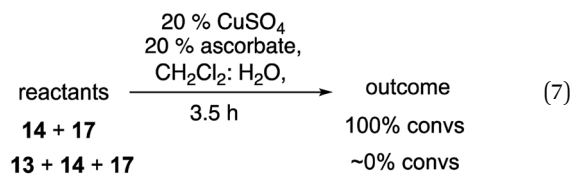
Table 5 CuAAC reaction of longer dialkyl peroxide



Alkyne, azide	X	Cu(I) ^a	Conv. (%)	Peroxide (yield, %)	Aldehyde (yield, %)
2, 16	H	CuI	87.6	22 (68%) 22i (6%)	24 (11%)
2, 16	H	CuSO ₄	90	22 (41%)	24 (37%)
4, 16	I	CuI	94	22i (76%) ^b	—
2, 16	H	CuI ^c	~10	22 (50%) 22i (3%)	—
2, 17	H	CuI	87	23 (55%) 23i (4%)	25 (29%)
2, 17	H	CuSO ₄	89	23 (39%)	25 (50%)

^a CuI/Et₃N (0.2 equiv.), THF; CuSO₄/sodium ascorbate, DCM/H₂O (biphasic). ^b Significant amount of unidentified polar byproduct. ^c Et₃SiH (1.9 equiv.) also present.

azide 17, which went to completion in just a few minutes under our typical reaction conditions, failed to proceed at all in the presence of a stoichiometric amount of peroxide 13 (eqn (7)).

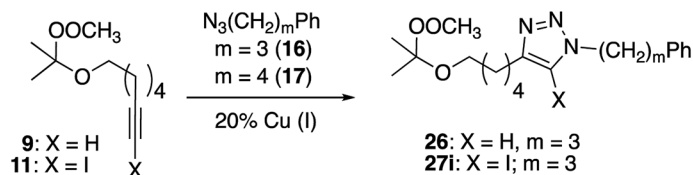


Nanoparticle functionalization

Having identified useful conditions for CuAAC reactions of alkynyl peroxides and peroxyacetals, we became interested in applying the chemistry to functionalization of nanoparticles. CuAAC-based functionalization of AuNPs, which exploits the high affinity of thiols for gold and the resulting ability to create azide-functionalized nanoparticles, has been widely applied to a variety of applications in both organic and aqueous media,^{11,14a,32–37} and we focused our attention on this system.

Azide-functionalized nanoparticles (**N₃Au**) were prepared using a variant of a reported procedure in which Au

Table 6 Click chemistry of hexynyl peroxyacetals



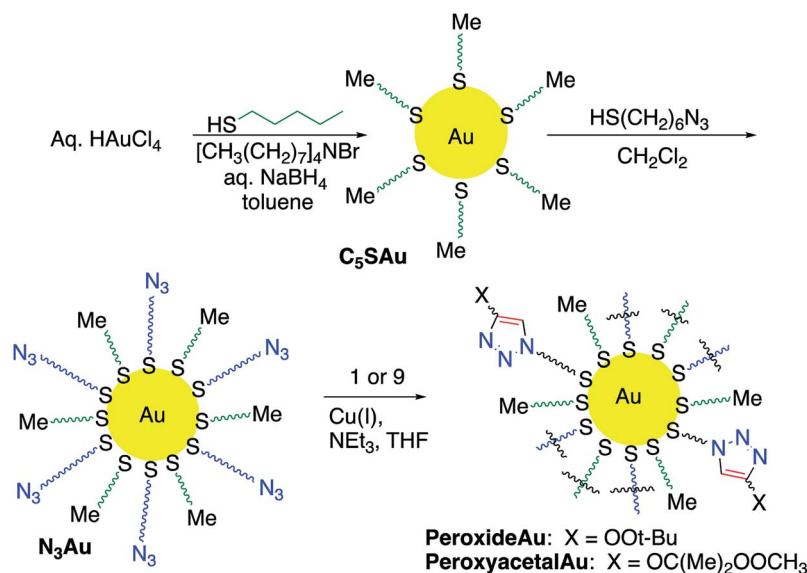
Alkyne	Azide	Cu(I) source	t (h)	Product (yield ^a , %)
9	17	CuI, Et ₃ N	3	26 (45%)
9	17	CuSO ₄ , ascorbate	24	Nr
11	16	CuI, Et ₃ N	18	27i (33%)

^a Isolated.



Table 7 CuAAC chemistry with peroxyalkyl azides

$ \begin{array}{ccc} t\text{-BuOO}(\text{CH}_2)_4\text{N}_3 & \xrightarrow{20\% \text{ Cu(I)}} & \text{OOt-Bu} \text{---} \text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{N} \text{---} \text{C}_5\text{H}_4\text{N} \text{---} (\text{CH}_2)_4\text{R} \\ \mathbf{12} & & \\ \text{R}(\text{CH}_2)_3\text{---} & & \\ \mathbf{14}: \text{R} = \text{Me} & & \mathbf{28}: \text{R} = \text{H} \text{ (34 \%)} \\ \mathbf{15}: \text{R} = \text{Ph} & & \mathbf{29}: \text{R} = \text{Ph} \text{ (15 \%)} \end{array} $				
Alkyne	Reagents	<i>t</i> (h)	Conv. (12)	Product (yield, %)
14	CuI/Et ₃ N	3	71%	28 (34%)
14	CuSO ₄ /Na ascorbate	24	80%	28 (29%)
15	CuI/Et ₃ N	3	76%	29 (15%)
15	CuSO ₄ /Na ascorbate	24	42%	29 (≤5%)



Scheme 4 Synthesis of peroxide-functionalized nanoparticles.

nanoparticles (AuNP) are subsequently reacted with a passivating and then a functionalized thiol (Scheme 4).³⁸ Addition of a slight excess of pentanethiol to a biphasic mixture of HAuCl₃ and tetraoctylammonium bromide resulted in a white suspension which reacted with freshly prepared aqueous sodium borohydride to generate a dark and opaque suspension of pentanethiolate-functionalized nanoparticles (C₅SAu). The partially passivated nanoparticles were isolated by centrifugation.^{11c} TEM (Fig. 2) spectra confirmed the presence of 2–5 nm particles. The presence of the alkylthio chains was evident by the ¹H signals for a terminal methyl group and the methylene groups at C₃/C₄ (ESI-Fig. 1†).³⁹ XPS analysis revealed the surface region of the nanoparticles to consist of 50.1% Au, 43.9% C, and 6.0% S, suggesting that approximately 75% of the Au surface remained accessible for further functionalization (ESI-Fig. 2–4†).⁴⁰ Repetition of the procedure using dodecanethiol, the preliminary passivating agent described in the original report, generated nanoparticles with similar properties (ESI-Fig. 5–8†).^{11c}

Treatment of the C₅SH-Au with excess 6-azido-hexane thiol furnished the azide-modified nanoparticle (N₃Au) which was

isolated and purified using centrifugation and resuspension. The presence of the azide was evident from the strong IR absorbance at 2094 cm^{−1} (Fig. 3, compare panels a and b) and

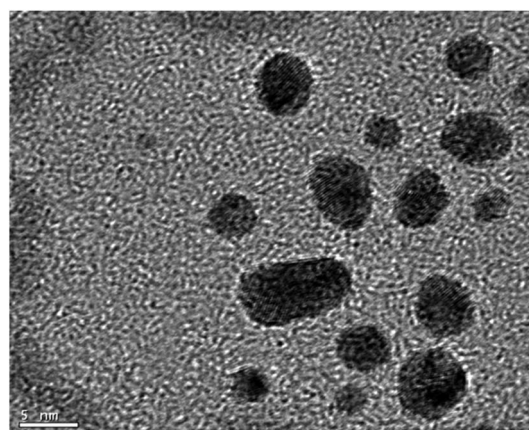


Fig. 2 TEM of C5SAu. Scale bar = 5 nm.



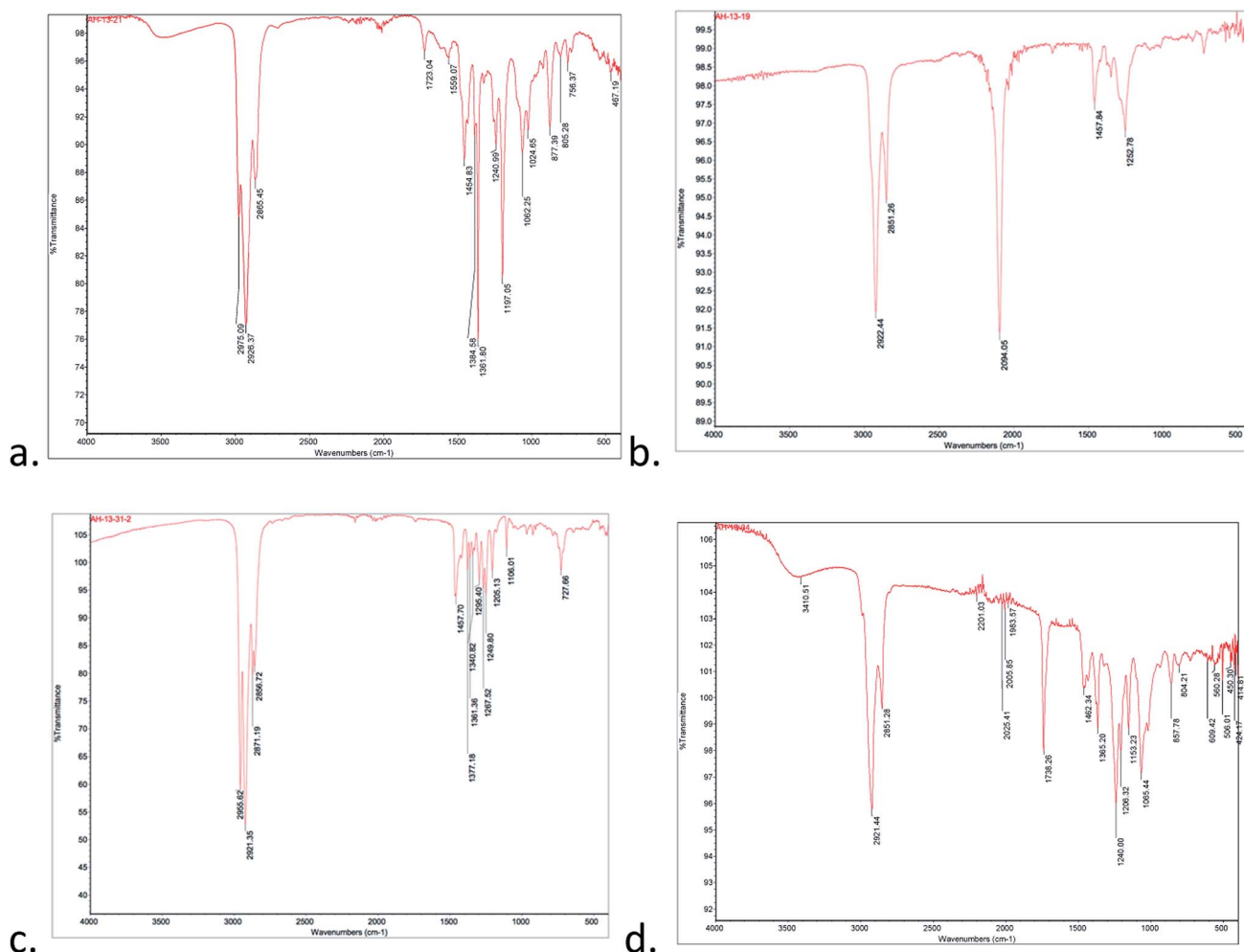


Fig. 3 IR of nanoparticles: (a) C₅SAu; (b) N₃Au; (c) after click with peroxyalkyne 1; (d) after click with alkynyl peroxyacetal 9.

by the new ¹H NMR signal at 3.31 ppm corresponding to CH₂N₃ (ESI-Fig. 9†); this signal was easily distinguished from signal for unbound azide (3.50 ppm), which was also observed in inadequately washed samples. The ~2 : 3 ratio of the integrals for the signals corresponding to CH₂N₃ vs. CH₂CH₃ suggests a similar extent of coverage. XPS suggested a surface composition of 54.0% Au, 34.2% C, 8.5% S, and 3.3% (ESI-Fig. 10–13†).

Peroxide functionalization of nanoparticles

Addition of triethylamine and cat. copper iodide into a THF solution containing peroxyalkyne 1 and N₃Au led to the immediate appearance of a new peroxide-active spot on TLC (see Experimental details).²⁶ After being quenched with aq. NH₄Cl, the reaction was concentrated and the residue purified away from residual reactant by repeated resuspension (MeOH) and centrifugation. The azide stretch was now absent in the IR spectrum (Fig. 3d), which instead showed weak tetrazole-related stretches at 1723 cm⁻¹ and 1559 cm⁻¹ and only minimal absorptions associated with an aldehyde (1705–1710 cm⁻¹). Evidence for the click reaction could also be seen in the ¹H NMR (ESI-Fig. 14†) as the loss of the δ 3.31 signal corresponding to the CH₂N₃ and new signals at 1.27 (*t*-Bu) and 3.99 (OO). XPS

established a product composition of 26.4% Au and 73.6% pentanethiol and peroxide triazole (ESI-Fig. 15–19†).

We also investigated click functionalization of azidoNP with alkynyl peroxy acetal 9. Consumption of the azide was obvious in the lack of the ~2100 cm⁻¹ IR stretch and the observation of a new C–H stretch at 2921 cm⁻¹ (Fig. 3d). Although ester formation was equally obvious as an IR stretch at 1738 cm⁻¹, the limited magnitude of this peak, combined with the comparison in the ¹H NMR (ESI-Fig. 20†) of signals for the MeOO (3.54 ppm) and the *gem*-dimethyl groups of the acetal (1.57 ppm) vs. the peak group at 3.86 ppm (CH₂O) of both acetal and ester, suggested the majority of the functionalized chain remained as peroxyacetal. XPS (ESI-Fig. 21–25†) suggested a surface composition of 20.0% Au and 80.0% pentanethiol and peroxyacetal triazole.

Discussion

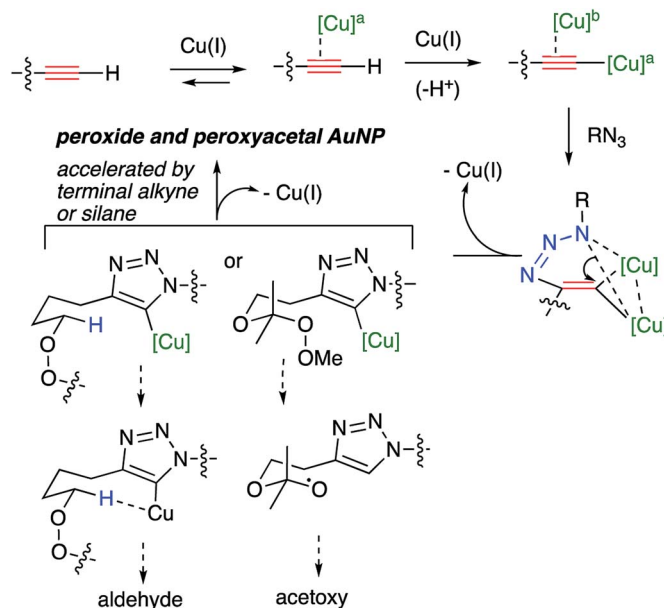
Our results demonstrate that both peroxides and peroxyacetals can, under suitable conditions, serve as modular components for Cu-promoted CuAAC click reactions. The click reactions of the alkynyl peroxides and peroxyacetals proceed at rates in line with literature reports, with reactions times of 0.5–1 h for

terminal alkynes and ≤ 0.75 h for 1-iodoalkynes.^{12,27} This is consistent with Zhu's observation that electron-poor alkynes are excellent substrates for CuAAC reactions.⁴¹ An "isomeric" set of click reactions involving simple alkynes and azidoalkyl peroxides is less efficient. The basis for inhibition of azide reactivity at a four-carbon distance from a peroxide remains unclear; however, we should note that we have observed a dampening influence of peroxides on S_N2 reactions across a similar span.⁴²

The use of CuI in organic solvents gives more consistent and often superior results compared to the use of CuSO_4 /ascorbate in biphasic media. The alkynyl peroxides and alkynyl peroxyacetal substrates, although stable towards Cu(I), are both prone to decomposition reactions in the presence of the reaction intermediates (Scheme 5). Interestingly, the efficiency of the reactions, the nature of the decomposition reactions, and the approaches to control of decomposition, are different for the two classes of substrates.

The dialkyl peroxides give modest to moderate yields under traditional CuAAC conditions due to the accumulation of a byproduct derived from fragmentation of the peroxide to an aldehyde. Aldehyde formation can be suppressed by use of an iodinated alkyne substrate, by performance of reactions in the presence of excess alkyne or added silane; under these conditions the yield of clicked peroxide is in the range of 62–68%. For the peroxyacetals, reactions proceed in lower yield, and, in the case of a shorter chain series (8), generate a significant amount of a byproduct in which the peroxyacetal has been fragmented to an acetoxy group. The formation of the ester is largely suppressed in a longer-chain peroxyacetal or for reactions of iodoalkynyl peroxyacetals. However, regardless of conditions, the yields obtained from peroxyacetals do not rival those obtained from the dialkyl peroxides.

An interpretation of our results in the context of the accepted mechanism of the CuAAC reaction (Scheme 6) suggests the decomposition reactions result from interaction of the peroxide or peroxyacetal with a heteroaryl organocopper intermediate generated at C_5 of the developing triazole.¹² Evidence in support of this hypothesis includes: the instability of the peroxyalkyne to any mixture containing Cu(I) and triazole; the lack of fragmentation observed for iodoalkynes; and the minimization of fragmentation the presence of excess alkyne or added triethylsilane, either of which would be expected to protonate the



Scheme 6 CuAAC mechanism and potential pathways for peroxide fragmentation.

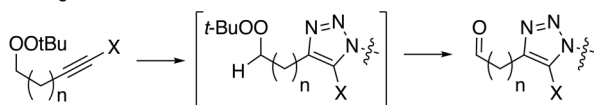
$\text{sp}^2\text{-RCu}$.³⁰ The clean formation of aldehyde, and the high barriers associated with electrochemical or chemical reduction of dialkyl peroxides,⁴³ suggest the decomposition is a heterolytic process involving abstraction of the adjacent C–H.^{1a}

The formation of an ester group (acetoxy) from fragmentations of the hydroperoxyacetals suggests the intermediacy of an alkoxy radical derived from cleavage of the peroxyacetal by an electron-donor; the likely reducing agent is the metalated triazole described above (Scheme 6).⁴⁴ In contrast to the reactions of dialkyl peroxides, byproduct formation for the peroxyacetals is not suppressed by the presence of excess alkyne or added silane; it is suppressed by use of iodoalkynyl substrates or by increasing the peroxide/alkyne distance.

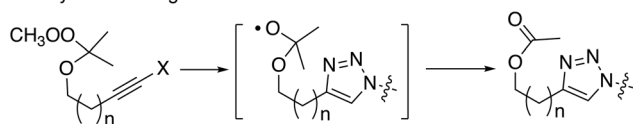
It is interesting to compare our results with reported CuAAC reactions of alkynyl derivatives of artemisinin-derived propargyl acetals and amides (Fig. 4); the approximate distance between the peroxide and the reaction site is similar in both systems.¹³ The heterolytic fragmentation we observed with dialkyl peroxides is not available in the artemisinin derivatives due to the lack of adjacent C–H groups. However, the fact that radical fragmentations are not observed in the artemisinin derivatives may indicate that the SET fragmentation requires close approach of the organocopper intermediates and the peroxyacetal, something that is possible in A and B and not possible in C or D.

CuAAC-based functionalization of AuNPs, which exploits the high affinity of thiols for gold has been widely applied to a variety of applications in both organic and aqueous media.^{11c,14b,c,32–37} At the outset of our studies, we were concerned by reports suggesting that CuAAC reactions on gold nanoparticles require longer reaction periods or use of microwave heating,^{45,46} and can be accompanied by nanoparticle aggregation and precipitation.^{45,47} Our results indicated that modular introduction of peroxides on nanoparticles *via* CuAAC suffers

Peroxides: side reaction observed for $n = 3$ or 5 ;
no fragmentation if $X = \text{I}$



Peroxyacetals: fragmentation observed for $n = 1$ but not $n = 3$



Scheme 5 Overview of decomposition processes.



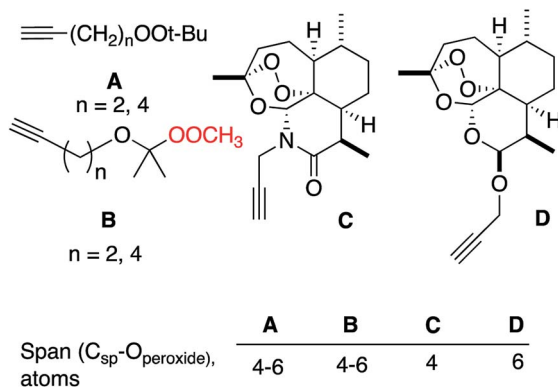


Fig. 4 Comparison with artemisinin-derived "click" substrates.

from no such limitations, and can be performed in less than eight hours and in the absence of microwave heating. Installation of dialkyl peroxides proceeds cleanly even in the absence of additives required in solution reactions, suggesting that the high local concentration of alkyne relative to the surface N_3 groups suppresses aldehyde formation. Peroxyacetals, which are much more activated than dialkyl peroxides towards cleavage,^{6,43} also undergo successful click reactions on surfaces. However, in this case, the side reaction observed in solution phase chemistry continued to be observed in reactions on nanoparticles. This reaction, in contrast to some results on the extremely hindered peroxyacetal core of artemisinin, is suppressed by the use of a substrate containing a greater span between the peroxide and alkyne groups.

With the exception of functionalized SAMs displaying peroxides (*via* ozonolysis of alkene-terminated monolayers)⁸ or diacyl peroxides (*via* condensation of H_2O_2 with carboxylic acids),^{48,49} there are few examples of covalent introduction of peroxide on surfaces. Much of this is likely to relate to concerns about compatibility of the surface with the functionalization method; for example, ozone is known to attack the thiol/Au interface and is incompatible with many electron-rich groups.^{18b,50} Our work demonstrates that the CuAAC reaction can be used to install both dialkyl peroxides and the more reactive peroxyacetals on nanoparticle surfaces.

Conclusions

Our work, the first systematic investigation of the factors influencing the use of click reactions for installing organic peroxides, demonstrates that the copper-assisted click offers a practical approach for modular introduction of dialkyl peroxides and peroxyacetals. The best results are obtained with dialkyl peroxides; a side reaction involving heterolytic fragmentation of the peroxide can be easily circumvented by choice of substrate or additive. Although installation of more reactive peroxyacetals can be complicated by a radical fragmentation, this process can be minimized by use of a greater span between the peroxide and the reaction site.

Our work also provides first example of the use of click chemistry for modular introduction of organic peroxides on

surfaces. Reactions are rapid, proceed to high conversion of nanoparticle surface functionality, and can be accomplished under mild conditions and without nanoparticle aggregation. Peroxide-functionalized surfaces could provide the basis for reactive or antimicrobial coatings,⁵¹ as well as a platform for investigating spatially-constrained interactions or reactions of peroxides or derived reactive oxygen species.⁵² Along these lines, the peroxyacetals investigated here, in addition to offering a source of alkoxy radicals, are known to be easily hydrolyzed to generate free hydroperoxides,¹⁶ while the simple *t*-alkyl peroxides, although nearly inert to simple reducing agents and $Fe(II)$, are known to be activated photochemically^{52,53} or by iron/thiol complexes.⁴³

Experimental procedures

General methods

All reagents and solvents were used as purchased except for CH_2Cl_2 (distilled from CaH_2), DMF (vacuum distilled from CaH_2), and THF (distilled from Na/Ph_2CO). All nonaqueous reactions were conducted under an atmosphere of N_2 in flame-dried glassware. Thin layer chromatography (TLC) was performed on 0.25 mm hard-layer silica plates. Developed plates were visualized by 254 nm UV lamp and/or by staining: 2.5% ammonium molybdate and 0.5% ceric sulfate in 10% aqueous sulfuric acid (general stain, after heating); 1% aq. potassium permanganate (alkynes); 1% *N,N*-dimethyl-*p*-phenylenediamine in 1 : 20 : 100 acetic acid/water/methanol (specific for peroxides; dialkyl peroxides and peroxyacetals can be visualized as a reddish or reddish-green spot upon heating);²⁶ or vanillin and sulfuric acid (3% each) in ethanol (general stain, after heating). Unless otherwise described, chromatography refers to silica flash chromatography.

NMR spectra were acquired in $CDCl_3$ at 400 (1H) or 100 MHz (^{13}C) unless otherwise noted. Chemical shifts are reported relative to residual chloroform (7.26 ppm, 1H ; 77.36 ppm, ^{13}C). 1H spectra are reported as chemical shift (multiplicity, *J* couplings in Hz, number of protons). IR spectra were recorded as neat films on a ZrSe crystal; selected absorbances are reported in cm^{-1} . High resolution mass spectra (HRMS) were obtained at the Nebraska Center for Mass Spectrometry at UNL. XPS and TEM spectra were acquired in the Nebraska Center for Nanoscience and Materials on a Thermo Scientific K-Alpha X-ray photoelectron spectrometer using a monochromated Al $K\alpha$ (1486.6 eV) X-ray source and a dual-beam flood source using a FEI Tecnai Osiris (scanning) transmission electron microscope.

CAUTION: Although we experienced no exotherms or explosions during the reported investigations, any work with peroxides of moderate or high active oxygen content, in particular any steps involving heating and/or concentration, should be conducted following standard precautions (adjusting scale to perceived hazard; concentration, and when necessary, reaction behind shields; some analysis of the thermal sensitivity of new products). Readers are directed towards a web-published overview of peroxide safety.^{10d}



Synthesis of peroxyalkynes

6-(*tert*-Butylperoxy)-1-hexyne (1) was prepared by an adaptation of published procedures.⁵⁴ To a solution of 6-chloro-1-hexyne (4.66 g, 40.0 mmol) in acetone (150 mL) was added sodium iodide (32.98 g, 220.0 mmol, 5.5 equiv.). The reaction was stirred under reflux for 18 hours. The cooled reaction was then diluted with water (50 mL) and extracted with hexanes (50 mL $\times$ 3). The combined organic layers were dried with Na₂SO₄ and the residue concentrated under reduced pressure to yield 7.5307 g (91%) of 6-iodo-1-hexyne as a yellow oil which was used without purification. Spectral details matched those previously reported.^{15,55} R_f = 0.42 (10% EA/Hex); ¹H δ 3.20 (t, J = 8.0, 2H), 2.22 (td, J = 3.5, J = 9.3, 2H), 1.92 (t, J = 9.3, 1H), 1.63 (p, J = 6.4, 2H); ¹³C δ 83.7, 69.0, 32.3, 29.2, 17.5, 6.2.

To a solution of CsOH monohydrate (4.50 g, 30.0 mmol, 1.2 equiv.) in DMF (100 mL) at 0 °C was added dropwise TBHP as an $\sim$ 5.5 M solution in decane (6.81 mL, 37.5 mmol, 1.5 equiv.). The mixture was stirred for 30 min, whereupon 6-iodo-1-hexyne (5.20 g, 25.0 mmol) was added. The reaction was allowed to slowly warm to room temperature. After 5 h, the reaction was quenched with water (30 mL) and extracted with hexanes (30 mL $\times$ 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure and the residue purified by column chromatography (5% EA/Hex) to yield 3.981 g (94%) of 6-(*tert*-butylperoxy)-1-hexyne as a colorless oil. Spectral details matched previous reports:^{54a} R_f = 0.56 (10% EA/Hex) ¹H δ 3.96 (t, J = 6.4, 2H), 2.22 (td, J = 7.0, J = 2.6, 2H), 1.94 (t, J = 2.6, 1H), 1.72 (m, 2H), 1.61 (m, 2H), 1.24 (s, 9H); ¹³C δ 84.3, 80.2, 74.5, 68.6, 27.2, 26.5, 25.4, 18.4; IR: 3314, 2924, 2119, 1362, 1198, 628; HRMS (ESI⁺, TOF) calcd for C₁₀H₁₈NaO₂ [M + Na]⁺: 193.1204; found: 193.1205.

8-(*tert*-Butylperoxy)-1-octyne (2) was prepared by a similar procedure as described above. Reaction of 8-chloro-1-octyne (3.3914 g, 23.45 mmol) and sodium iodide (19.366 g, 129.2 mmol, 5.5 equiv.) furnish 5.0135 g (91%) of 8-iodo-1-octyne as a yellow oil which was used without further purification. Spectral details matched those previously reported:⁵⁶ R_f : 0.43 (10% EA/Hex); ¹H δ 3.19 (t, J = 7.0, 2H), 2.19 (td, J = 6.6, J = 2.2, 2H), 1.94 (t, J = 2.6, 1H), 1.83 (p, J = 6.7, 2H), 1.53 (m, 2H), 1.41–1.43 (m, 4H); ¹³C δ 84.5, 68.4, 33.5, 30.1, 28.3, 27.7, 18.4, 7.1.

By the same method as employed for synthesis of **1**, reaction of the iodo-octyne (3.8046 g, 16.50 mmol) with CsOH monohydrate (3.310 g, 22.1 mmol, 1.3 equiv.) and TBHP as an $\sim$ 5.5 M solution in decane (4.50 mL, 24.75 mmol, 1.5 equiv., $\sim$ 5.5 M solution in decane) in DMF, furnished 3.0118 g (92%) of 8-(*tert*-butylperoxy)-1-octyne (**2**) as a colorless oil: R_f : 0.55 (10% EA/Hex); ¹H δ 3.93 (t, J = 6.6, 2H), 2.18 (td, J = 6.9, J = 2.5, 2H), 1.93 (t, J = 2.6, 1H), 1.50–1.64 (m, 4H), 1.35–1.46 (m, 4H), 1.24 (s, 9H); ¹³C δ 84.7, 80.2, 75.1, 68.3, 28.7, 28.5, 27.9, 26.4, 25.8, 18.4 IR: 3312, 2945 (s), 2956, 1456, 1360 (s), 1198 (s); HRMS (ESI⁺, TOF) calcd for C₁₂H₂₂NaO₂ [M + Na]⁺: 221.1517; found: 221.1510.

1-Iodo-6-*tert*-butylperoxy hexyne (3) was prepared using an adaptation of a published procedure.¹⁷ To a solution of 6-(*tert*-butylperoxy)-1-hexyne (850 mg, 4.99 mmol) in MeOH (20 mL) was added KI (1.02 g, 6.13 mmol, 1.2 equiv.), followed by TBHP

as an $\sim$ 5.5 M solution in decane (1.36 mL, 7.48 mmol, 1.2 equiv.). The reaction was stirred at room temperature for 18 hours. The reaction was quenched with saturated aqueous Na₂S₂O₃ (10 mL), washed with brine (10 mL), and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na₂SO₄, and the residue concentrated under reduced pressure to yield 1.1083 g (75%) of iodoalkynyl peroxide **3** as a light yellow oil: R_f = 0.55 (20% EA/Hex); ¹H δ 3.95 (t, J = 6.3, 2H), 2.40 (t, J = 6.8, 2H), 1.70 (p, J = 6.1, 2H), 1.59 (p, J = 7.0, 2H), 1.24 (s, 9H) ¹³C δ 94.4, 80.3, 74.4, 27.2, 26.5, 25.4, 20.8, –6.8; IR 2925, 1362, 1196, 877; HRMS (ESI⁺, TOF) calcd for C₁₀H₁₇INaO₂ [M + Na]⁺: 319.0171; found: 319.0174.

1-Iodo-8-*tert*-butylperoxy octyne (4). By a similar procedure as employed for synthesis of iodoalkynyl peroxide **3**, reaction of 8-(*tert*-butylperoxy)-1-octyne (714.9 mg, 3.60 mmol), KI (685.5 mg, 4.13 mmol, 1.2 equiv.), and TBHP (1.0 mL, 5.5 mmol, 1.5 equiv.) furnished, after column chromatography (10% EA/Hex), 901.1 mg (77%) of iodo-octynyl peroxide **4** as a light yellow oil: R_f : 0.57 (20% EA/Hex); ¹H δ 3.95 (t, J = 6.7, 2H), 2.38 (t, J = 7.0, 2H), 1.62 (p, J = 7.1, 2H), 1.54 (p, J = 7.0, 2H), 1.26–1.36 (m, 4H), 1.26 (s, 9H); ¹³C δ 94.8, 80.2, 75.0, 28.7, 28.5, 27.9, 26.5, 25.8, 20.9, –7.4; IR: 2977, 2937, 2860, 1362, 1197; HRMS (ESI⁺, TOF) calcd for C₁₂H₂₁INaO₂ [M + Na]⁺: 347.0484; found: 347.0485.

Synthesis of alkynyl peroxyacetals

3-((2-Hydroperoxypropan-2-yl)oxy)-1-propyne (5) and **3-((2-hydroperoxypropan-2-yl)peroxy)propan-2-yl)oxy)-1-propyne (5a)** were prepared using a procedure reported for unsaturated alcohols.¹⁹ To a solution of propargyl alcohol (0.97 mL, 16.9 mmol, 3 equiv.) in CH₂Cl₂ (50 mL) was added 2,3-dimethyl-2-butene (0.65 mL, 5.47 mmol). This mixture was cooled to –78 °C whereupon a gaseous stream of O₃/O₂ was admitted (pipette, $\sim$ 0.5–1 mmol O₃/minute). When the reaction was complete (TLC) introduction of O₃/O₂ was halted and N₂ was briefly bubbled through the reaction mixture. The solution was carefully concentrated under reduced pressure (CAUTION) and the residue purified by column chromatography (5% ether/pentane) to yield 210.8 mg (20%) of hydroperoxyperoxide **5a** as a colorless oil, followed by traces of volatile hydroperoxyacetal **5**. (CAUTION: the nearly 16% active oxygen content of hydroperoxyperoxide **5a** suggests a moderate to high potential for exothermic decomposition.) As a result, this preparation was confined to <20 mmol scale and products were handled carefully (see general Discussion on safety, above).^{10d}

5a: R_f = 0.46 (20% EA/Hex); ¹H δ 9.61 (s, 1H), 4.31 (t, J = 2.4, 2H), 2.43 (t, J = 2.3, 1H), 1.51 (s, 6H), 1.42 (s, 6H); ¹³C δ 109.4, 106.8, 80.4, 74.3, 51.2, 23.5, 20.9; IR: 3362, 3286, 2999, 2946, 1369, 1176, 1034, 826, 620; HRMS (ESI⁺, TOF) calcd for C₉H₁₆NaO₅ [M + Na]⁺: 227.0895; found: 227.0896.

5: (trace byproduct; incompletely characterized as a colorless oil): R_f = 0.43 (20% EA/Hex); ¹H δ 9.88 (s, 1H), 4.20 (t, J = 2.4, 2H), 2.45 (t, J = 2.4, 1H), 1.41 (s, 6H); ¹³C δ 105.9, 81.3, 73.7, 49.9, 22.7.

4-((2-Hydroperoxypropan-2-yl)oxy)-1-butyne (6) and **4-((2-hydroperoxypropan-2-yl)peroxy)propan-2-yl)oxy)-1-butyne (6a)** were prepared using the same procedure described above. To



a solution of 3-butyne-1-ol (1.97 mL, 26.0 mmol, 2 equiv.) in CH_2Cl_2 (125 mL) was added 2,3-dimethyl-2-butene (1.56 mL, 13.1 mmol). This mixture was cooled to -78°C whereupon a gaseous stream of O_3/O_2 was admitted (approximately 1 mmol O_3 /minute). When the reaction as judged complete, by TLC, introduction of O_3/O_2 was halted and N_2 was briefly bubbled through the reaction mixture. The solution was carefully concentrated under reduced pressure and the residue (CAUTION) purified by column chromatography (5% ether/pentane) to yield 455.6 mg (24%) of **6** as an colorless oil, accompanied by 141.5 mg (5%) of **6a**. CAUTION: The products shown here are hydroperoxides or peroxy hydroperoxides with active oxygen contents suggesting a moderate to high potential for exothermic decomposition. These reactions were deliberately run on a modest scale and products should be handled carefully.

6a: $R_f = 0.46$ (20% EA/Hex); ^1H δ 9.75 (s, 1H), 3.80 (t, $J = 9.6$, 2H), 2.59 (td, $J = 6.1$, $J = 3.6$, 2H), 2.01 (t, $J = 3.6$, 1H), 1.53 (s, 6H), 1.47 (s, 6H); ^{13}C δ 109.4, 106.8, 80.4, 74.3, 51.2, 23.5, 20.9; IR: 3293, 2997, 2946, 1367, 1196, 1177, 1141, 1045, 834, 636; HRMS (ESI^+ , TOF) calcd for $\text{C}_9\text{H}_{16}\text{NaO}_5$ [$\text{M} + \text{Na}$] $^+$: 241.1052; found: 241.1052.

6: $R_f = 0.53$ (20% EA/Hex); ^1H δ 8.46 (s, 1H), 3.65 (t, $J = 5.4$, 2H), 2.50 (td, $J = 6.0$, $J = 2.6$, 2H), 2.09 (t, $J = 2.6$, 1H), 1.49 (s, 6H); ^{13}C δ 105.7, 83.7, 69.8, 60.0, 22.6, 20.5; IR: 3427, 3287, 2995, 2947, 1367, 1201, 1158, 1051, 816, 639; HRMS (ESI^+ , TOF) calcd for $\text{C}_7\text{H}_{12}\text{NaO}_3$ [$\text{M} + \text{Na}$] $^+$: 167.0684; found: 167.0681.

6-((2-Hydroperoxypropan-2-yl)oxy)-1-hexyne (7). By a similar procedure as described above, a mixture of 5-hexyn-1-ol (0.88 mL, 7.98 mmol, 3 equiv.) in CH_2Cl_2 (30 mL) and 2,3-dimethyl-2-butene (0.32 mL, 2.69 mmol) were reacted with O_3/O_2 to generate, after careful concentration and chromatography 328.6 mg (71%) of hydroperoxyacetal **7** as a colorless oil: $R_f = 0.41$ (20% EA/Hex); ^1H δ 7.89 (s, 1H), 3.51 (t, $J = 6.5$, 2H), 2.22 (td, $J = 6.7$, $J = 2.6$, 2H), 1.95 (t, $J = 2.6$, 1H), 1.73–1.57 (m, 4H), 1.39 (s, 6H); ^{13}C δ 105.3, 84.5, 68.7, 61.2, 29.1, 25.3, 22.8, 18.3; IR: 3296, 2993, 2943, 2872, 2117, 1740, 1366, 1199, 1151, 1061, 866, 629; HRMS (ESI^+ , TOF) calcd for $\text{C}_9\text{H}_{16}\text{NaO}_3$ [$\text{M} + \text{Na}$] $^+$: 195.0997; found: 195.0995.

4-((2-Methylperoxypropan-2-yl)oxy)-1-butyne (8) was prepared using an adaptation of a reported procedure.^{21,57} To a solution of hydroperoxyacetal **6** (359 mg, 2.49 mmol) in THF (10 mL) was added potassium *tert*-butoxide (350 mg, 3.12 mmol, 1.2 equiv.), followed by methyl iodide (0.19 mL, 3.05 mmol, 1.2 equiv.). After the reaction had stirred at room temperature for 5 minutes, it was quenched with water (10 mL) and the resulting mixture extracted with ether (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 and concentrated under reduced pressure. The residue was purified by column chromatography (10% ether/pentane) to yield 262.2 mg (66%) of peroxyacetal **8** as a colorless oil: $R_f = 0.54$ (20% EA/Hex); ^1H δ 3.85 (s, 3H), 3.66 (t, $J = 9.8$, 2H), 2.47 (td, $J = 9.8$, $J = 3.6$, 2H), 1.97 (t, $J = 3.6$, 1H), 1.40 (s, 6H); ^{13}C δ 104.9, 81.5, 69.4, 63.4, 60.2, 23.2, 20.3; IR: 2992, 2943, 2894, 1379, 1200, 1152, 1042, 864; HRMS (ESI^+ , TOF) calcd for $\text{C}_8\text{H}_{14}\text{NaO}_3$ [$\text{M} + \text{Na}$] $^+$: 181.0814; found: 181.0841.

6-((2-Methylperoxypropan-2-yl)oxy)-1-hexyne (9). By a similar procedure as described for above, a solution of hydroperoxyacetal **7** (146 mg, 0.85 mmol) in THF (5 mL) was reacted with potassium *tert*-butoxide (114 mg, 1.02 mmol, 1.2 equiv.) and methyl iodide (0.10 mL, 1 mmol, 1.2 equiv.), to furnish, after purification, 142.5 mg (90%) of peroxyacetal **9** as an colorless oil: $R_f = 0.45$ (20% EA/Hex); ^1H δ 3.84 (s, 3H), 3.53 (t, $J = 8.3$, 2H), 2.23 (td, $J = 9.3$, $J = 3.4$, 2H), 1.93 (t, $J = 3.1$, 1H), 1.71–1.59 (m, 4H), 1.38 (s, 6H); ^{13}C δ 104.7, 84.6, 68.4, 63.3, 61.0, 29.1, 25.4, 23.5, 23.2, 18.3; IR: 3294, 2943, 2117, 1365, 1205, 1153, 1067, 1015, 869, 629; HRMS (ESI^+ , TOF) calcd for $\text{C}_{10}\text{H}_{18}\text{NaO}_3$ [$\text{M} + \text{Na}$] $^+$: 209.1154; found: 209.1153.

1-Iodo-4-((2-methylperoxypropan-2-yl)oxy)-1-butyne (10). By a similar procedure as employed for synthesis of iodoalkynyl peroxide **3**, reaction of peroxyacetal **8** (141 mg, 0.89 mmol) in MeOH (3 mL) with KI (183 mg, 1.10 mmol, 1.2 equiv.) followed by TBHP as an ~ 5.5 M solution in decane (0.25 mL, 1.38 mmol, 1.2 equiv.) furnished, after concentration under reduced pressure. 206.7 mg (81%) of iodoalkynyl peroxyacetal **10** as a light yellow oil: $R_f = 0.54$ (10% EA/Hex); ^1H δ 3.85 (s, 3H), 3.64 (t, $J = 7.3$, 2H), 2.64 (t, $J = 7.4$, 2H), 1.39 (s, 6H); ^{13}C δ 104.9, 91.5, 63.4, 60.2, 23.2, 22.6, -5.4 ; IR: 2992, 2940, 2891, 1366, 1199, 1152, 1073, 1013, 862; HRMS (ESI^+ , TOF) calcd for $\text{C}_8\text{H}_{13}\text{INaO}_2$ [$\text{M} + \text{Na}$] $^+$: 306.9807; found: 306.9811.

1-Iodo-6-((2-methylperoxypropan-2-yl)oxy)-1-hexyne (11). By a similar procedure as employed for synthesis of **3**, reaction of peroxyacetal **9** (242 mg, 1.30 mmol) in MeOH (10 mL) with KI (267 mg, 1.61 mmol, 1.2 equiv.) and TBHP as an ~ 5.5 M solution in decane (0.36 mL, 1.98 mmol, 1.5 equiv.) furnished iodoalkynyl peroxyacetal **11** (400.8 mg, 98%) of as a light yellow oil: $R_f = 0.49$ (10% EA/Hex); ^1H δ 3.85 (s, 3H), 3.52 (t, $J = 6.5$, 2H), 2.40 (t, $J = 6.8$, 2H), 1.69–1.54 (m, 4H), 1.38 (s, 6H); ^{13}C δ 104.7, 94.7, 63.3, 61.0, 29.2, 25.4, 23.2, 20.8, -7.1 ; IR: 3299, 2940, 1466, 1378, 1365, 1204, 1152, 1013, 858, 632; HRMS (ESI^+ , TOF) calcd for $\text{C}_{10}\text{H}_{17}\text{INaO}_2$ [$\text{M} + \text{Na}$] $^+$: 335.0120; found: 335.0126.

Synthesis of peroxy azides

1-Azido-4-((tert-butylperoxy))butane (12). To a solution of 1-bromo-4-chlorobutane (2.57 g, 15.0 mmol) in DMF (40 mL) was added sodium azide (1.96 mg, 30.1 mmol, 2 equiv.). The reaction was stirred at room temperature for 24 hours. The reaction was diluted with (20 mL) water and extracted with hexanes (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 and the residue concentrated under reduced pressure to yield 1.8975 g (95%) of 1-azido-4-chlorobutane as a colorless oil which was used without further purification: $R_f = 0.49$ (10% EA/Hex). Spectral details matched those previously reported.²² $R_f = 0.49$ (10% EA/Hex).

A solution of the 1-azido-4-chlorobutane (2.40 g, 18.0 mmol) and sodium iodide (15.0 g, 100 mmol, 5.5 equiv.) was refluxed in acetone (150 mL) for 18 hours. The reaction was then diluted with water (60 mL) and extracted with hexanes (50 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 and the residue concentrated under reduced pressure to yield 3.6943 g (91%) of 1-azido-4-iodobutane as an orange oil which was used without further purification. Spectral details matched those



previously reported.²² $R_f = 0.5$ (10% EA/Hex); ^1H δ 3.32 (t, $J = 6.6$, 2H), 3.20 (t, $J = 6.7$, 2H), 1.91 (p, $J = 7.4$, 2H), 1.71 (p, $J = 7.4$, 2H); ^{13}C δ 50.5, 30.6, 30.0, 7.2.

By a similar procedure as applied for synthesis of peroxide **1**, reaction of CsOH monohydrate (127 mg, 0.85 mmol, 1.2 equiv.) in DMF (12 mL) at 0 °C, TBHP as an ~5.5 M solution in decane (0.19 mL, 1.05 mmol, 1.5 equiv., ~5.5 M solution in decane), and 1-azido-4-iodobutane (158 mg, 0.70 mmol) furnished, after purification by chromatography (5% EA/Hex), 80.9 mg (62%) of the azidobutyl *t*-butyl peroxide (**12**) as a colorless oil (CAUTION: low molecular weight azide/peroxide): $R_f = 0.51$ (10% EA/Hex); ^1H δ 3.96 (m, 2H), 3.30 (t, $J = 6.4$, 2H), 1.71–1.66 (m, 4H), 1.24 (s, 9H); ^{13}C δ 80.3, 74.7, 51.4, 26.5, 26.0, 25.4; IR: 2977, 2931, 2093, 1456, 1362, 1245, 1196, 881; HRMS (ESI⁺, TOF) calcd for $\text{C}_8\text{H}_{17}\text{N}_3\text{NaO}_2$ [$\text{M} + \text{Na}$]⁺: 210.1218; found: 210.1213.

((6-Azidoheptan-2-yl)peroxy)triethylsilane (13). Into a RBF containing CH_2Cl_2 (300 mL), were added sequentially iodine (7.61 g, 30.0 mmol), PPh_3 (7.87 g, 30.0 mmol) and imidazole (2.04 g, 30.0 mmol). The mixture was stirred for 15 minutes at 0 °C, followed by the addition of 5-hexenol (2.50 g, 25.0 mmol). Reaction was stirred at room temperature and filtered through silica with hexanes (600 mL). The filtrate was concentrated under reduced pressure to yield 3.8557 g (75%) of 6-iodohexene as a light orange oil which was used without further purification. Spectral details matched those previously reported.⁵⁸ $R_f = 0.58$ (10% EA/Hex); ^1H δ 5.02 (dq, $J = 17.0$, $J = 1.8$, 1H), 4.96 (ddt, $J = 10.0$, 1.8, 0.9, 1H), 5.78 (dddd, $J = 17.2$, $J = 10.5$, $J = 6.7$, $J = 6.7$, 1H), 3.19 (t, $J = 7.0$, 2H), 2.08 (q, $J = 7.0$, 2H), 1.84 (p, $J = 7.5$, 2H), 1.50 (p, $J = 7.5$, 2H); ^{13}C δ 138.2, 115.1, 33.0, 32.7, 29.8, 7.0.

To a solution of the iodoheptene (1.20 g, 5.71 mmol) in DMF (20 mL) was added sodium azide (1.11 g, 17.1 mmol, 3 equiv.). The reaction was stirred at room temperature for 18 hours and then diluted with water (10 mL). The combined hexane extracts (3 $\times$ 15 mL) were dried with Na_2SO_4 and the residue concentrated under reduced pressure to yield 688.5 mg (98%) of 6-azido-heptene as a colorless oil which was used without further purification. Spectral details matched those previously reported:⁵⁹ $R_f = 0.59$ (10% EA/Hex); ^1H δ 5.01 (dq, $J = 17.2$, $J = 1.6$, 1H), 4.97 (d, $J = 10.2$, 1H), 5.79 (dddd, $J = 17.0$, $J = 10.5$, $J = 6.7$, $J = 6.7$, 1H), 3.27 (t, $J = 6.8$, 2H), 2.09 (q, $J = 7.1$, 2H), 1.62 (p, $J = 7.5$, 2H), 1.48 (p, $J = 7.6$, 2H); ^{13}C δ 138.3a 115.1, 51.5, 33.3, 28.4, 26.0.

To a solution of the azidoheptene (113 mg, 0.90 mmol) in EtOH (10 mL) was added triethylsilane (0.29 mL, 1.82 mmol, 2 equiv.), followed by $\text{Co}(\text{acac})_2$ (25.7 mg, 0.10 mmol, 0.1 equiv.). The reaction was stirred at room temperature under O_2 for 18 hours.²³ The solution was concentrated under reduced pressure and the residue purified by column chromatography (5% EA/Hex) to yield 114.0 mg (46%) of azido peroxide **13** as a colorless oil: $R_f = 0.5$ (10% EA/Hex); ^1H δ 4.01 (s, $J = 5.8$, 1H), 3.27 (t, $J = 6.9$, 2H), 1.61–1.40 (m, 6H), 1.20 (d, $J = 6.2$, 3H), 0.99 (t, $J = 7.9$, 9H), 0.69 (q, $J = 8.0$, 6H); ^{13}C δ 81.2, 51.5, 34.0, 29.1, 22.8, 18.5, 6.9, 3.9; IR: 2938, 2877, 2093, 1459, 1242, 1006, 796, 727; HRMS (ESI⁺, TOF) calcd for $\text{C}_{12}\text{H}_{27}\text{N}_3\text{NaO}_2\text{Si}$ [$\text{M} + \text{Na}$]⁺: 296.1770; found: 296.1772.

Hexyne (14) was used as received

Hex-5-yn-1-yl benzene (15) (100848-88-2). Alkyne **15** was prepared using a variant of a reported procedure.⁶⁰ To a solution of ethynyl trimethylsilane (998.3 mg, 3.84 mmol) in THF (10 mL) was added *n*-BuLi (2.7 mL, 4.32 mmol, 1.1 equiv.) at –78 °C, followed by 4-(iodobutyl)benzene (386.2 mg, 3.93 mmol, 1.02 equiv.). The reaction was allowed to warm to room temperature and stirred for 4 hours. The reaction was quenched with water (30 mL) and the combined ethyl acetate extracts (3 $\times$ 30 mL) were dried with Na_2SO_4 . Concentration under reduced pressure yielded 825.2 mg (85%) of trimethyl(6-phenylhex-1-yn-1-yl)silane (2253948-27-3) as a colorless oil which was used directly in the following reaction. $R_f = 0.60$ (10% EA/Hex). Spectra details matched those in a literature report.⁶¹

To a solution of the crude alkynyl silane (743.2 mg, 3.09 mmol) in methanol (25 mL) was added potassium hydroxide (304.0 mg, 5.42 mmol, 1.75 equiv.). The reaction was stirred at room temperature for 2 hours.⁶² The reaction was quenched with water (40 mL) and extracted with hexanes (30 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (5% EA/Hex) to yield 499.0 mg (89%) of the hex-5-yn-1-yl benzene as a colorless oil. Spectral details matched those previously reported:^{61,63} $R_f = 0.65$ (10% EA/Hex); ^1H δ 7.31 (m, 2H), 7.24–7.21 (m, 3H), 2.67 (t, $J = 7.9$, 2H), 2.25 (td, $J = 2.6$, $J = 7.1$, 2H), 1.98 (t, $J = 2.6$, 1H), 1.78 (p, $J = 7.8$, 2H), 1.61 (p, 7.5, 2H); ^{13}C δ 142.4, 128.5, 128.4, 125.9, 84.5, 68.4, 35.5, 30.5, 28.1, 18.4; IR: 3305, 2927, 2854, 1495, 1453, 744, 696.

3-Azidopropylbenzene (16). Iodine (11.74 g, 46.3 mmol), PPh_3 (11.89 g, 45.3 mmol) and imidazole (3.16 g, 46.4 mmol) were sequentially added to a solution of CH_2Cl_2 (300 mL). The mixture was stirred for 15 minutes at 0 °C, whereupon 3-phenyl-1-propanol (5.07 g, 37.2 mmol) was added. The reaction was stirred at room temperature for 5 hours and filtered through a ~3" pad of silica with hexanes (600 mL). The filtrate was concentrated under reduced pressure to yield 9.05 g (98%) of 3-iodopropylbenzene as a light orange oil which was used without further purification. Spectral details matched those previously reported.⁶⁴ $R_f = 0.50$ (10% EA/Hex); ^1H δ 7.33 (t, $J = 7.4$, 2H), 7.26 (t, $J = 8.1$, 1H), 7.24 (d, $J = 6.6$, 2H), 3.21 (t, $J = 6.8$, 2H), 2.77 (t, $J = 7.3$, 2H), 2.17 (p, $J = 7.2$, 2H); ^{13}C δ 140.5, 128.7, 128.6, 126.3, 36.4, 35.0, 6.5.

To a solution of the iodopropyl benzene (1.10 g, 4.48 mmol) in DMF (20 mL) was added sodium azide (875.6 g, 13.4 mmol, 3 equiv.). The reaction was stirred at room temperature for 18 hours. The reaction was diluted with water (10 mL) and extracted with CH_2Cl_2 (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 and the residue concentrated under reduced pressure to yield 694.1 mg (96%) of 3-azidopropylbenzene as a colorless oil. Spectral details matched those previously reported.⁶⁵ $R_f = 0.52$ (10% EA/Hex); ^1H δ 7.34 (t, $J = 7.4$, 2H), 7.26 (t, $J = 7.4$, 1H), 7.22 (d, $J = 7.3$, 2H), 3.32 (t, $J = 6.8$, 2H), 2.75 (t, $J = 7.5$, 2H), 1.95 (p, $J = 7.2$, 2H); ^{13}C δ 141.0, 128.6, 128.6, 126.3, 50.8, 32.9, 30.6.



4-Azidobutylbenzene (17). Using a similar procedure as described above, reaction of a solution of iodine (7.18 g, 28.3 mmol), PPh₃ (7.40 g, 28.2 mmol), imidazole (1.91 g, 28.1 mmol) and 4-phenyl-1-butanol (3.47 g, 23.1 mmol) in CH₂Cl₂ (300 mL) furnished 4.6282 g (77%) of 4-iodobutylbenzene as a light orange oil. Spectral details matched those previously reported.⁶⁶ $R_f = 0.51$ (10% EA/Hex); ^1H δ 7.32 (dt, $J = 5.4, J = 1.5$, 2H), 7.24 (dt, $J = 8.4, 1.2$, 1H), 7.21 (d, $J = 6.8$, 2H), 3.23 (t, $J = 6.8$, 2H), 2.67 (t, $J = 7.6$, 2H), 1.89 (m, 2H), 1.77 (m, 2H); ^{13}C δ 141.9, 128.5, 126.0, 34.9, 33.1, 32.3, 6.9.

Using a similar procedure as described above, reaction of the 4-iodobutylbenzene (1.25 g, 4.81 mmol) with sodium azide (969 mg, 14.9 mmol, 3 equiv.) in DMF (20 mL) furnished 811.7 mg (97%) of 4-azidobutylbenzene (17) as a colorless oil. Spectral details matched those previously reported.⁶⁶ $R_f = 0.53$ (10% EA/Hex); ^1H δ 7.32 (t, $J = 7.3$, 2H), 7.22 (t, $J = 7.3$, 1H), 7.21 (d, $J = 7.3$, 2H), 3.31 (t, $J = 6.5$, 2H), 2.68 (t, $J = 7.6$, 2H), 1.75 (m, 2H), 1.67 (m, 2H); ^{13}C δ 142.0, 128.5, 128.5, 126.0, 51.5, 35.5, 28.6.

6-Azidohehexane-1-thiol (18). This molecule, which is commercially available, has been previously described without characterization.⁶⁷

Synthesis of *S*-(6-hydroxyhexyl)ethanethioate was adapted from a reported procedure.⁶⁷ To a solution of 6-bromohexanol (3.07 g, 17.0 mmol) in acetone (175 mL) was added potassium thioacetate (3.92 g, 34.3 mmol, 2 equiv.). The reaction was stirred at 40 °C for 18 hours. The reaction was then quenched with 1 M NaHCO₃ (50 mL) and the mixture extracted with ether (50 mL $\times$ 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure, and the residue was purified by column chromatography (5–20% ether/pentane) to yield 1.89 g (63%) of the thioester as a red oil. Spectral details matched those previously reported.⁶⁸ $R_f = 0.30$ (30% EA/Hex); ^1H δ 3.62 (t, $J = 6.4$, 2H), 2.86 (t, $J = 6.8$, 2H), 2.31 (s, 3H), 1.53–1.62 (m, 4H), 1.35–1.43 (m, 4H); ^{13}C δ 196.2, 62.9, 32.7, 30.8, 29.6, 29.1, 28.6, 25.3; IR: 3292, 2944, 1380, 1205, 1157, 1064, 831, 630.

The following reagents were sequentially added to CH₂Cl₂ (300 mL): iodine (2.86 g, 11.3 mmol, 1.2 equiv.), PPh₃ (2.96 g, 11.3 mmol, 1.2 equiv.) and imidazole (769.3 mg, 11.3 mmol, 1.2 equiv.). The mixture was stirred for 15 minutes at 0 °C, whereupon *S*-(6-hydroxyhexyl)ethanethioate (1.66 g, 9.42 mmol) was added. The reaction was stirred at room temperature for 5 hours and filtered through a 7.5 cm pad of silica with hexanes (600 mL). The filtrate was concentrated under reduced pressure to yield 1.98 g (73%) of *S*-(6-iodohexyl)ethanethioate as a brown oil which was used directly for the next reaction: $R_f = 0.40$ (5% EA/Hex); ^1H δ 3.20 (t, $J = 7.0$, 2H), 2.89 (t, $J = 7.3$, 2H), 2.35 (s, 3H), 1.84 (p, $J = 6.8$, 2H), 1.61 (p, $J = 6.9$, 2H), 1.37–1.47 (m, 4H); ^{13}C δ 196.1, 33.4, 30.8, 30.1, 29.4, 29.1, 27.8, 7.0; IR: 2929, 1686, 1132, 954, 623; HRMS (ESI⁺, TOF) calcd for C₈H₁₅IOS [M]⁺: 285.9888; found: 285.9898.

To a solution of *S*-(6-azidohehexyl)ethanethioate (1.16 g, 4.05 mmol) in DMF (20 mL) was added sodium azide (1.45 g, 22.3 mmol, 5.5 equiv.). The reaction was stirred at room temperature for 18 hours. The reaction was diluted with water (10 mL) and extracted with CH₂Cl₂ (15 mL $\times$ 3). The combined

organic layers were dried with Na₂SO₄ and the residue concentrated under reduced pressure to yield 762.6 mg (94%) of azido/ethioester as an orange oil which was used without further purification. Spectral details matched those previously reported.⁶⁹ $R_f = 0.32$ (5% EA/Hex); ^1H δ 3.25 (t, $J = 7.1$, 2H), 2.86 (t, $J = 7.1$, 2H), 2.32 (s, 3H), 1.54–1.61 (m, 4H), 1.37–1.40 (m, 4H); ^{13}C δ 196.0, 51.5, 30.8, 29.5, 29.0, 28.8, 28.4, 26.3; IR: 2933, 2090, 1688, 1256, 1132, 951, 624.

The deblocking of the thioacetate employed a procedure reported for a nine-carbon homolog.⁷⁰ To a solution of *S*-(6-azidohehexyl)ethanethioate (762.6 g, 3.79 mmol) in MeOH (70 mL) was added conc. aq. HCl (4.0 mL). The reaction was stirred at reflux for 3 hours and then allowed to cool. The reaction was diluted with water (20 mL) and extracted with diethyl ether (25 mL $\times$ 3). The combined organic layers were dried with Na₂SO₄ and the residue concentrated under reduced pressure to yield 559.8 mg (93%) of 6-azidohehexane-1-thiol as a strong-smelling orange oil. $R_f = 0.35$ (5% EA/Hex); ^1H δ 3.27 (t, $J = 6.6$, 2H), 2.53 (q, $J = 7.2$, 2H), 1.57–1.64 (m, 4H), 1.37–1.43 (m, 4H); IR: 3930, 3857, 2400 (w); 2086 (s); 1454, 1253; ^{13}C δ 51.5, 33.9, 28.9, 28.0, 26.3, 24.6 IR: 2930, 2087, 1254; HRMS (ESI⁺, TOF) calcd for C₆H₁₃N₃S [M]⁺: 159.0830; found: 159.0867, 316.1512 (disulfide).

General procedure for click reactions

Click reactions were conducted based upon adaptations of a reported procedure.^{25b}

General procedure for click reactions with CuI. To a solution of alkyne (1 equiv.) and azide (1 equiv.) in THF was added CuI (0.2 equiv.), followed by triethylamine (3 equiv.). The reaction was stirred at room temperature for 3 hours, then diluted with water and extracted with ethyl acetate ($\times$ 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure and the residue purified by column chromatography.

General procedure for click reactions with CuSO₄. To a solution of alkyne (1 equiv.) and azide **16** (1 equiv.) in CH₂Cl₂ was added an aqueous solution containing CuSO₄ (0.2 equiv.) and sodium ascorbate (0.2 equiv.). After stirring at rt for 3 h, the reaction was diluted with water and extracted with ethyl acetate ($\times$ 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure and the residue purified by column chromatography.

4-(4-(*tert*-Butylperoxy)butyl)-1-(3-phenylpropyl)-1H-1,2,3-triazole (19). To a solution of peroxyhexyne **1** (203 mg, 1.19 mmol) and azide **16** (193 mg, 1.20 mmol) in THF (5 mL) was added CuI (47.6 mg, 0.25 mmol, 0.2 equiv.),^{25b} followed by triethylamine (0.50 mL, 3.6 mmol, 3 equiv.) and triethylsilane (0.21 mL, 1.3 mmol, 1.1 equiv.). The reaction was stirred at room temperature for 24 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 156.2 mg (38%) of triazole **19** as a yellow oil: $R_f = 0.28$ (40% EA/Hex); ^1H δ 7.29 (t, $J = 7.4$, 2H), 7.26 (s, 1H), 7.23 (t, $J = 7.4$, 1H), 7.19 (d, $J = 7.1$, 2H), 4.33 (t, $J =$



7.1, 2H), 3.99 (t, $J = 6.3$, 2H), 2.77 (t, $J = 7.4$, 2H), 2.67 (t, $J = 7.5$, 2H), 2.25 (p, $J = 7.3$, 2H), 1.78 (m, 2H), 1.69 (m, 2H), 1.25 (s, 9H); ^{13}C δ 148.0, 140.4, 128.7, 128.5, 126.4, 120.7, 80.2, 74.8, 49.5, 32.7, 31.8, 27.6, 26.5, 26.2, 25.6; IR: 2976, 2932, 2866, 1497, 1454, 1362, 1196, 1046, 878, 746, 670, 493; HRMS (ESI⁺, TOF) calcd for $\text{C}_{19}\text{H}_{29}\text{N}_3\text{NaO}_2$ $[\text{M} + \text{Na}]^+$: 354.2157; found: 354.2155.

4-(4-(*tert*-Butylperoxy)butyl)-1-(4-phenylbutyl)-1H-1,2,3-triazole (20) and 4-(4-(*tert*-butylperoxy)butyl)-5-iodo-1-(4-phenylbutyl)-1H-1,2,3-triazole (20i). To a solution of 6-(*tert*-butylperoxy)-1-hexyne (238 mg, 1.40 mmol) and 4-azidobutylbenzene (209 mg, 1.19 mmol) in THF (5 mL) was added CuI (47.6 mg, 0.25 mmol, 0.2 equiv.), followed by triethylamine (0.60 mL, 4.29 mmol, 3 equiv.).⁴⁶ The reaction was stirred at room temperature for 1 hour. The reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 154.4 mg (37%) of triazole **20** as a yellow oil. Reactions conducted in the presence of CuI often contained traces of the corresponding 5-iodotriazole (**20i**).

20: $R_f = 0.27$ (40% EA/Hex); ^1H δ 7.29 (t, $J = 7.4$, 2H), 7.24 (s, 1H), 7.21 (t, $J = 7.4$, 1H), 7.16 (d, $J = 7.2$, 2H), 4.33 (t, $J = 7.2$, 2H), 3.98 (t, $J = 6.1$, 2H), 2.75 (t, $J = 7.2$, 2H), 2.66 (t, $J = 7.5$, 2H), 1.93 (p, $J = 7.5$, 2H), 1.80–1.62 (m, 6H), 1.25 (s, 9H); ^{13}C δ 150.0, 141.6, 128.5, 128.5, 126.1, 120.6, 80.2, 74.8, 50.1, 35.3, 29.9, 28.3, 27.6, 26.4, 26.2, 25.6 ppm; IR: 3026, 2934, 1454, 1362, 1196, 1044, 747, 699; HRMS (ESI⁺, TOF) calcd for $\text{C}_{20}\text{H}_{31}\text{N}_3\text{NaO}_2$ $[\text{M} + \text{Na}]^+$: 368.2314; found: 368.2312.

20i: $R_f = 0.31$ (40% EA/Hex); ^1H δ 7.21 (t, $J = 7.2$, 2H), 7.18 (m, 3H), 4.37 (t, $J = 7.2$, 2H), 3.99 (t, $J = 6.4$, 2H), 2.70 (t, $J = 7.2$, 2H), 2.68 (t, $J \sim 7$, 2H), 1.95 (p, $J = 7.2$, 2H), 1.80 (m, 2H), 1.76 (m, 4H), 1.25 (s, 9H); ^{13}C δ 151.60, 141.7, 128.58, 128.55, 126.1, 120.6, 80.3, 78.1, 74.8, 50.6, 35.3, 29.5, 28.2, 27.6, 26.4, 26.1, 25.8 ppm; IR: 2976, 2936, 1454, 1362, 1197, 1044, 743, 699; HRMS (ESI⁺, TOF) calcd for $\text{C}_{20}\text{H}_{30}\text{IN}_3\text{NaO}_2$ $[\text{M} + \text{Na}]^+$: 494.1280; found: 494.1276.

4-(1-(4-Phenylbutyl)-1H-1,2,3-triazol-4-yl)butanal (21). This colorless oil, which coelutes with the parent peroxide, was a minor byproduct in many of the click reactions but became a major byproduct for reactions conducted for long periods or when the peroxy triazole products were resubjected to reaction conditions: $R_f = 0.27$ (40% EA/Hex); ^1H δ 9.79 (t, $J = 0.7$, 1H), 7.32–7.16 (m, 6H), 4.34 (t, $J = 7.2$, 2H), 2.78 (t, $J = 7.4$, 2H), 2.68 (t, $J = 7.5$, 2H), 2.55 (td, $J = 1.4$, $J = 7.3$, 2H), 2.04 (p, $J = 7.3$, 2H), 1.94 (p, $J = 7.6$, 2H), 1.67 (m, 2H); ^{13}C (176 MHz) δ 202.2, 147.2, 141.6, 128.6, 128.5, 126.2, 120.8, 50.2, 43.3, 35.3, 32.1, 28.3, 27.2, 21.9; IR: 3135, 3025, 2925, 2723, 1722, 1454, 1275, 748, 701; HRMS (ESI⁺, TOF) calcd for $\text{C}_{20}\text{H}_{31}\text{N}_3\text{NaO}_2$ $[\text{M} + \text{Na}]^+$: 294.1582; found: 294.1588.

Stability of peroxyalkyne towards CuI

To 6-(*tert*-butylperoxy)-1-hexyne **1** (170 mg, 1.0 mmol) in THF (5 mL) was added CuI (38 mg, 0.20 mmol, 0.2 equiv.). The reaction was stirred at room temperature for 24 hours. The reaction was quenched with water and extracted with ethyl acetate (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 ,

concentrated under reduced pressure to yield 153.4 mg (90%) of recovered starting material as a colorless oil: $R_f = 0.55$ (10% EA/Hex).

Reactions in the presence of TBTA

CuSO₄. To 6-(*tert*-butylperoxy)-1-hexyne **1** (272 mg, 1.6 mmol) and 4-phenyl-butyl-azide **17** (228 mg, 1.3 mmol) in 1 : 1 DMSO/water (10 mL) was added CuSO_4 (25.5 mg, 0.16 mmol, 0.1 equiv.), sodium ascorbate (47.5 mg, 0.24 mmol, 0.15 equiv.), and TBTA (8.4 mg, 0.016 mmol, 0.01 equiv.). The reaction was stirred at room temperature for 1.5 hours. The reaction was quenched with water and extracted with ethyl acetate (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and purified by column chromatography (10% EA/Hex) to yield 89.9 mg (20%) of triazole **20** as a yellow oil: $R_f = 0.28$ (40% EA/Hex).

CuI. To 6-(*tert*-butylperoxy)-1-hexyne **1** (238 mg, 1.4 mmol) and azide **17** (245 mg, 1.4 mmol) in DMSO : water (2.5 mL : 2.5 mL) was added CuI (57 mg, 0.30 mmol, 0.2 equiv.) and triethylamine (0.6 mL, 4.2 mmol, 3 equiv.), followed by TBTA (7.4 mg, 0.014 mmol, 0.01 equiv.). The reaction was stirred at room temperature for 3 hours. The reaction was quenched with water and extracted with ethyl acetate (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and purified by column chromatography (10% EA/Hex) to yield 166.3 mg (35%) of triazole **20** as a yellow oil: $R_f = 0.28$ (40% EA/Hex).

Reaction of iodoalkyne

4-(4-(*tert*-Butylperoxy)butyl)-5-iodo-1-(3-phenylpropyl)-1H-1,2,3-triazole (19i). To a solution of 1-iodo-6-(*tert*-butylperoxy)hexyne **3** (118 mg, 0.40 mmol) and 3-azidopropylbenzene **16** (64.5 mg, 0.41 mmol) in THF (3.5 mL) was added CuI (7.6 mg, 0.04 mmol, 0.1 equiv.), followed by triethylamine (0.12 mL, 0.86 mmol, 2 equiv.).¹² The reaction was stirred at room temperature for 24 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 116.4 mg (64%) of iodotetrazole **19i** as a yellow oil: $R_f = 0.36$ (20% EA/Hex); ^1H δ 7.32 (t, $J = 7.5$, 2H), 7.24 (t, $J = 7.5$, 1H), 7.22 (d, $J = 7.6$, 2H), 4.37 (t, $J = 7.3$, 2H), 3.99 (t, $J = 6.4$, 2H), 2.70 (td, $J = 7.5$, $J = 3.4$, 4H), 2.25 (p, $J = 7.4$, 2H), 1.80 (p, $J = 7.6$, 2H), 1.68 (p, $J = 6.3$, 2H), 1.25 (s, 9H); ^{13}C δ 151.6, 140.4, 128.7, 128.6, 126.4, 80.2, 78.1, 74.7, 50.1, 32.6, 31.3, 27.6, 26.5, 26.0, 25.8; IR: 3031, 2936, 2098, 1461, 1361, 1196 1022, 880, 752, 699, 509; HRMS (ESI⁺, TOF) calcd for $\text{C}_{19}\text{H}_{28}\text{IN}_3\text{NaO}_2$ $[\text{M} + \text{Na}]^+$: 480.1124; found: 480.1126.

Competition of iodoalkynyl peroxide with simple alkyne

To a solution of 6-(*tert*-butylperoxy)-1-iodohexyne **3** (0.5 mmol), 5-hexyn-1-ol (0.5 mmol) and 3-phenylpropyl azide **16** (0.5 mmol) in THF (5 mL) was added CuI (0.050 mmol, 0.1 equiv.), followed by triethylamine (1.0 mmol, 2 equiv.). The reaction was stirred at room temperature for 18 hours. The reaction was quenched with water and extracted with ethyl acetate (15 mL $\times$ 3). The



combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and purified by column chromatography (10% EA/Hex) to yield 64.5 mg (29%) of 4-(4-(*tert*-butylperoxy)butyl)-5-iodo-1-(3-phenylpropyl)-1*H*-1,2,3-triazole (**19i**) as a yellow oil.

4-(6-(*tert*-Butylperoxy)hexyl)-1-(3-phenylpropyl)-1*H*-1,2,3-triazole (6/3**) (**22**) and 4-(6-(*tert*-butylperoxy)hexyl)-5-iodo-1-(3-phenylpropyl)-1*H*-1,2,3-triazole and 6-(1-(3-phenylpropyl)-1*H*-1,2,3-triazol-4-yl)hexanal (**24**).** To a solution of 8-(*tert*-butylperoxy)-1-octyne **2** (304.1 mg, 1.53 mmol) and 3-azidopropylbenzene **16** (268.4 mg, 1.66 mmol, 1.1 equiv.) in THF (6 mL) was added CuI (64.8 mg, 0.25 mmol, 0.2 equiv.), followed by triethylamine (0.65 mL, 4.66 mmol, 3 equiv.).³ The reaction was stirred at room temperature for 3 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 374.3 mg (68%) of peroxy triazole **22**, 45.9 mg (6%) of iodotriazole **22i** and 47.1 mg (11%) of aldehyde **24**.

22: R_f : 0.28 (40% EA/Hex); ^1H δ 7.31 (t, J = 7.5, 2H), 7.24 (s, 1H), 7.22 (t, J = 7.6, 1H), 7.19 (d, J = 7.4, 2H), 4.33 (t, J = 7.1, 2H), 3.94 (t, J = 6.5, 2H), 2.73 (t, J = 7.6, 2H), 2.66 (t, J = 7.5, 2H), 2.25 (p, J = 7.3, 2H), 1.69 (p, J = 7.4, 2H), 1.61 (p, J = 6.8, 2H), 1.36–1.45 (m, 4H), 1.25 (s, 9H); ^{13}C δ 148.4, 140.4, 128.7, 128.5, 126.4, 120.6, 80.2, 75.1, 49.4, 32.7, 31.8, 29.5, 29.2, 27.9, 26.5, 26.1, 25.7; IR: 2976, 2933, 2859, 1454, 1361, 1197, 1047; HRMS (ESI⁺, TOF) calcd for $\text{C}_{21}\text{H}_{33}\text{N}_3\text{NaO}_2$ [$\text{M} + \text{Na}$]⁺: 382.2470; found: 382.2467.

22i: R_f : 0.32 (40% EA/Hex); ^1H δ 7.31 (t, J = 7.3, 2H), 7.21–7.24 (m, 3H), 4.36 (t, J = 7.2, 2H), 3.94 (t, J = 6.6, 2H), 2.68 (app p, J = 8.1, J = 7.6, 4H), 2.25 (p, J = 7.2, 2H), 1.71 (p, J = 7.1, 2H), 1.61 (p, J = 6.6, 2H), 1.40–1.41 (m, 4H), 1.25 (s, 9H); ^{13}C δ 151.9, 140.4, 128.7, 128.6, 126.4, 80.1, 78.0, 75.1, 50.1, 32.6, 31.3, 29.1, 29.0, 27.9, 26.5, 26.1, 26.1; IR: 2974, 2934, 2861, 1453, 1362; HRMS (ESI⁺, TOF) calcd for $\text{C}_{21}\text{H}_{32}\text{IN}_3\text{NaO}_2$ [$\text{M} + \text{Na}$]⁺: 508.1437; found: 508.1423.

24: R_f : 0.27 (40% EA/Hex); ^1H δ 9.76 (t, J = 1.69, 1H), 7.30 (t, J = 7.4, 2H), 7.25 (s, 1H), 7.22 (t, J = 7.2, 1H), 7.18 (d, J = 7.3, 2H), 4.32 (t, J = 7.2, 2H), 2.73 (t, J = 7.6, 2H), 2.65 (t, J = 7.6, 2H), 2.44 (td, J = 1.7, J = 7.3, 2H), 2.24 (p, J = 7.6, 2H), 1.69 (m, 2H), 1.58 (p, J = 6.8, 2H) 1.42 (m, 2H); ^{13}C (176 MHz) δ 202.7, 148.4, 140.4, 128.7, 128.5, 126.4, 120.6, 62.9, 49.5, 43.9, 32.7, 31.8, 25.5, 21.9, 14.3; IR: 2928, 2856, 1720, 1453, 1212, 1047, 1029 cm^{-1} ; HRMS (ESI⁺, TOF) calcd for $\text{C}_{17}\text{H}_{23}\text{N}_3\text{NaO}$ [$\text{M} + \text{Na}$]⁺: 308.1739; found: 308.1734.

4-(6-(*tert*-Butylperoxy)hexyl)-1-(3-phenylpropyl)-1*H*-1,2,3-triazole (22**) via CuSO_4 .** To a solution of 8-(*tert*-butylperoxy)-1-octyne **2** (301.6 mg, 1.52 mmol) and 3-azidopropylbenzene **16** (260.5 mg, 1.62 mmol, 1.1 equiv.) in CH_2Cl_2 (4 mL) was added 2 mL of an aqueous solution containing CuSO_4 (50.7 mg, 0.32 mmol, 0.2 equiv.) and sodium ascorbate (65.4 mg, 0.33 mmol, 0.2 equiv.).^{12,29} After stirring at rt for 3 h, the reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 223.9 mg (41%) of 4-(6-(*tert*-butylperoxy)hexyl)-1-(3-phenylpropyl)-1*H*-

1,2,3-triazole (**22**) accompanied by 139.0 mg (37%) of aldehyde **24**.

4-(6-(*tert*-Butylperoxy)hexyl)-1-(3-phenylpropyl)-1*H*-1,2,3-triazole (22**) in presence of silane.** To a solution of 8-(*tert*-butylperoxy)-1-octyne **2** (263.8 mg, 1.33 mmol) and 3-azidopropylbenzene **16** (261.0 mg, 1.34 mmol, 1.0 equiv.) in THF (5 mL) was added CuI (53.1 mg, 0.28 mmol, 0.2 equiv.),¹² followed by triethylamine (0.60 mL, 4.3 mmol, 3 equiv.) and triethylsilane (0.45 mL, 2.50 mmol, 1.9 equiv.). The reaction was stirred at room temperature for 24 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 239.0 mg (50%) of 4-(6-(*tert*-butylperoxy)hexyl)-1-(3-phenylpropyl)-1*H*-1,2,3-triazole as a yellow oil (spectra described previously).

4-(6-(*tert*-Butylperoxy)hexyl)-5-iodo-1-(3-phenylpropyl)-1*H*-1,2,3-triazole (22i**) from iodoalkyne **3**.** To a solution of 1-iodo-8-*tert*-butylperoxy octyne **4** (297.2 mg, 0.917 mmol) and 3-azidopropylbenzene **16** (159.0 mg, 0.99 mmol, 1.1 equiv.) in THF (6 mL) was added CuI (28.3 mg, 0.15 mmol, 0.2 equiv.), followed by triethylamine (0.30 mL, 2.15 mmol, 2 equiv.).^{27c} The reaction was stirred at room temperature for 24 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 338.4 mg (76%) of iodotriazole **22i** as a yellow oil (described previously).

4-(6-(*tert*-Butylperoxy)hexyl)-1-(4-phenylbutyl)-1*H*-1,2,3-triazole (23**), 4-(6-(*tert*-butylperoxy)hexyl)-5-iodo-1-(4-phenylbutyl)-1*H*-1,2,3-triazole (**23i**) and 6-(1-(4-phenylbutyl)-1*H*-1,2,3-triazol-4-yl)hexanal (**25**).** To a solution of 8-(*tert*-butylperoxy)-1-octyne **2** (305.4 mg, 1.54 mmol) and 4-azidobutylbenzene **17** (277.0 mg, 1.58 mmol, 1.0 equiv.) in THF (6 mL) was added CuI (64.2 mg, 0.34 mmol, 0.2 equiv.), followed by triethylamine (0.65 mL, 4.66 mmol, 3 equiv.). The reaction was stirred at room temperature for 24 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL $\times$ 3). The combined organic layers were dried with Na_2SO_4 , concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 315.7 mg (55%) of 29.8 mg (4%) of iodotriazole **23i**, 315.7 mg (55%) of triazole **23** and 134.6 mg (45%) of aldehyde **25**.

23: R_f : 0.28 (40% EA/Hex); ^1H δ 7.30 (t, J = 7.3, 2H), 7.22 (s, 1H), 7.21 (t, J = 7.5, 1H), 7.16 (d, J = 7.4, 2H), 4.33 (t, J = 7.2, 2H), 3.94 (t, J = 6.6, 2H), 2.72 (t, J = 7.7, 2H), 2.67 (t, J = 7.5, 2H), 1.94 (p, J = 7.6, 2H), 1.58–1.72 (m, 6H), 1.40–1.41 (m, 4H), 1.26 (s, 9H); ^{13}C δ 148.4, 141.6, 128.5, 128.5, 126.1, 120.5, 80.2, 75.1, 50.1, 35.3, 29.9, 29.5, 29.2, 28.3, 27.9, 26.5, 26.1, 25.7; IR: 2977, 2920, 2860, 1454, 1362; HRMS (ESI⁺, TOF) calcd for $\text{C}_{22}\text{H}_{35}\text{N}_3\text{NaO}_2$ [$\text{M} + \text{Na}$]⁺: 396.2627; found: 396.2612.

23i:²⁷ R_f : 0.32 (40% EA/Hex) ^1H δ 7.30 (t, J = 6.8, 2H), 7.22 (t, J = 7.2, 1H), 7.18 (d, J = 7.0, 2H), 4.37 (t, J = 7.3, 2H), 3.95 (t, J = 6.3, 2H), 2.64–2.70 (m, 4H), 1.96 (p, J = 7.0, 2H), 1.68–1.73 (m, 4H), 1.62 (p, J = 6.5, 2H), 1.38–1.46 (m, 4H), 1.26 (s, 9H); ^{13}C δ 151.9, 141.6, 128.5, 128.5, 126.1, 80.2, 77.9, 75.1, 50.6, 35.2, 29.8, 29.5, 29.1, 29.0, 28.1, 27.9, 26.5, 26.1; IR: 2088, 2920, 2858,



1454, 1361; HRMS (ESI⁺, TOF) calcd for C₂₂H₃₄IN₃NaO₂ [M + Na]⁺: 522.1593; found: 522.1581.

25: R_f: 0.27 (40% EA/Hex); ¹H δ 9.77 (t, J = 1.7, 1H), 7.30 (t, J = 7.4, 2H), 7.22 (s, 1H), 7.21 (t, J = 7.4, 1H), 7.16 (d, J = 7.4, 2H), 4.33 (t, J = 7.2, 2H), 2.73 (t, J = 7.6, 2H), 2.67 (t, J = 7.6, 2H), 1.94 (p, J = 7.7, 2H), 1.63–1.73 (m, 4H), 1.59 (p, J = 6.6, 2H), 1.37–1.45 (m, 4H); ¹³C (176 MHz) δ 202.8, 148.4, 141.6, 128.5, 128.5, 126.1, 120.5, 62.9, 50.1, 43.9, 35.3, 32.7, 29.9, 28.3, 25.5, 21.9; IR: 2920, 2858, 1721, 1454; HRMS (ESI⁺, TOF) calcd for C₁₈H₂₅N₃NaO [M + Na]⁺: 322.1895; found: 322.1882.

4-(6-(*tert*-Butylperoxy)hexyl)-1-(4-phenylbutyl)-1H-1,2,3-triazole (23) and 6-(1-(4-phenylbutyl)-1H-1,2,3-triazol-4-yl)hexanal (25). To a solution of 8-(*tert*-butylperoxy)-1-octyne **2** (305.1 mg, 1.54 mmol) and 4-azidobutylbenzene **17** (272.3 mg, 1.55 mmol, 1.0 equiv.) in CH₂Cl₂ (4 mL) was added a solution of CuSO₄ (59.2 mg, 0.37 mmol, 0.2 equiv.) and sodium ascorbate (75.5 mg, 0.38 mmol, 0.2 equiv.) in water (2 mL).¹² The reaction was stirred at room temperature for 24 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL × 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 225.7 mg (39%) of peroxy triazole **23** as a yellow oil accompanied by 231.4 mg (50%) of aldehyde **25**.

Click reaction of peroxyacetals

To a solution of 4-((2-methylperoxypropan-2-yl)oxy)-1-butyne **8** (166 mg, 1.05 mmol) and 4-azidobutylbenzene **17** (193 mg, 1.10 mmol) in THF (6 mL) was added CuI (38.1 mg, 0.20 mmol, 0.2 equiv.), followed by triethylamine (0.28 mL, 2.01 mmol, 3 equiv.). The reaction was stirred at room temperature for 1.5 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL × 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield a light yellow oil as a nearly inseparable mixture of 4-2-((2-methylperoxy)propan-2-yl)oxyethyl)-1-(4-phenylbutyl)-1H-1,2,3-triazole and 4-(2-acetoxyethyl)-1-(4-phenylbutyl)-1H-1,2,3-triazole. The peroxyacetal, typically formed in less than 10% yield, was contaminated with variable amounts of the ester.

4-(2-((2-(Methylperoxy)propan-2-yl)oxy)ethyl)-1-(4-phenylbutyl)-1H-1,2,3-triazole. R_f = 0.34 (30% EA/Hex); ¹H δ 7.32 (t, J = 7.1, 2H), 7.27 (s, 1H), 7.24 (t, J = 7.7, 1H), 7.23 (d, J = 7.5, 2H), 4.38 (t, J = 7.2, 2H), 3.86 (t, J = 7.6, 2H), 3.84 (s, 3H), 2.97 (t, J = 7.5, 2H), 2.71 (t, J = 7.3, 2H), 2.26 (p, J = 7.4, 2H), 1.28 (s, 6H); ¹³C δ 149.3, 140.4, 128.7, 128.6, 126.4, 104.9, 63.3, 60.7, 50.2, 32.6, 31.4, 29.8, 27.4, 23.2.

4-(4-((2-(Methylperoxy)propan-2-yl)oxy)butyl)-1-(3-phenylpropyl)-1H-1,2,3-triazole (26). To a solution of 6-((2-methylperoxypropan-2-yl)oxy)-1-hexyne **10** (102 mg, 0.55 mmol) and 4-azidobutylbenzene **17** (98.1 mg, 0.56 mmol) in THF (5 mL) was added CuI (20.9 mg, 0.11 mmol, 0.2 equiv.), followed by triethylamine (0.23 mL, 1.7 mmol, 3 equiv.). The reaction was stirred at room temperature for 3 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL × 3). The combined organic layers were dried with Na₂SO₄,

concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 88.5 mg (45%) of the peroxyacetal/triazole **26** as a yellow oil: R_f = 0.28 (30% EA/Hex); ¹H δ 7.32 (t, J = 7.3, 2H), 7.27 (s, 1H), 7.23 (t, J = 7.4, 1H), 7.19 (d, J = 7.2, 2H), 4.33 (t, J = 7.1, 2H), 3.85 (s, 3H), 3.56 (t, J = 6.5, 2H), 2.77 (t, J = 7.5, 2H), 2.67 (t, J = 7.5, 2H), 2.25 (p, J = 7.3, 2H), 1.82–1.56 (m, 4H), 1.40 (s, 6H); ¹³C δ 148.0, 140.4, 128.7, 128.6, 126.4, 120.7, 104.7, 63.3, 61.3, 49.5, 32.7, 31.8, 29.6, 26.2, 25.6, 23.2; IR: 3025, 2926, 2858, 2202, 1736, 1454, 1365, 1207, 1066, 748, 701; HRMS (ESI⁺, TOF) calcd for C₁₉H₂₉N₃NaO₃₀. [M + Na]⁺: 370.2107; found: 370.2103.

5-Iodo-4-(4-((2-(methylperoxy)propan-2-yl)oxy)butyl)-1-(3-phenylpropyl)-1H-1,2,3-triazole (27i). To a solution of 1-iodo-6-((2-methylperoxypropan-2-yl)oxy)-1-hexyne **11** (172 mg, 0.55 mmol) and 3-azidopropylbenzene **16** (96.7 mg, 0.60 mmol) in THF (5 mL) was added CuI (10.5 mg, 0.055 mmol, 0.1 equiv.), followed by triethylamine (0.15 mL, 1.08 mmol, 2 equiv.). The reaction was stirred at room temperature for 24 hours. The reaction was diluted with water (10 mL) and extracted with EA (15 mL × 3). The combined organic layers were dried with Na₂SO₄, concentrated under reduced pressure and the residue purified by column chromatography (10% EA/Hex) to yield 84.5 mg (33%) of peroxyacetal/iodotriazole **27i** as a light yellow oil: R_f = 0.31 (30% EA/Hex); ¹H δ 7.31 (t, J = 7.0, 2H), 7.23 (t, J = 7.0, 1H), 7.21 (d, J = 7.7), 4.36 (t, J = 7.2, 2H), 3.85 (s, 3H), 3.56 (t, J = 6.6, 2H), 2.70 (t, J = 7.6, 2H), 2.25 (p, J = 7.4, 2H), 1.84–1.62 (m, 4H), 1.40 (s, 6H); ¹³C δ 151.7, 140.4, 128.6, 128.5, 126.4, 104.7, 78.1, 63.3, 61.3, 50.1, 32.6, 31.3, 29.5, 26.0, 25.7, 23.2; IR: 3356, 3026, 2934, 2860, 1733, 1453, 1210, 1028, 745, 699; HRMS (ESI⁺, TOF) calcd for C₁₅H₂₀IN₃NaO [(M – C₄H₈O₂) + Na]⁺; elim. with loss of peroxyacetal]⁺: 408.0549; found: 408.0549.

Click reaction with azidoalkyl peroxides

4-Butyl-1-(4-(*tert*-butylperoxy)butyl)-1H-1,2,3-triazole (28). To a solution of 1-azido-4-(*tert*-butylperoxy)butane **12** (142 mg, 0.76 mmol) and 1-hexyne **14** (65.7 mg, 0.80 mmol) in THF (3.5 mL) was added CuI (30.5 mg, 0.16 mmol, 0.2 equiv.), followed by triethylamine (0.32 mL, 2.30 mmol, 3 equiv.). After stirring for 3.5 hours, the reaction was diluted with water (10 mL). The combined EA extracts (15 mL × 3) were dried with Na₂SO₄, and concentrated under reduced pressure. The residue was purified by column chromatography (10% EA/Hex) to yield 69.5 mg (34%) of peroxy triazole **28** as a light yellow oil: R_f = 0.25 (40% EA/Hex); ¹H δ 7.26 (s, 1H), 4.34 (t, J = 7.2, 2H), 3.95 (t, J = 6.2, 2H), 2.70 (t, J = 7.7, 2H), 1.98 (p, J = 7.5, 2H), 1.64 (p, J = 7.4, 2H), 1.62 (p, J = 7.4, 2H), 1.37 (tt, J = 14.9, J = 7.4, 2H), 1.22 (s, 9H), 0.92 (t, J = 7.3, 3H); ¹³C δ 148.5, 120.6, 80.4, 74.2, 50.0, 31.7, 27.5, 26.4, 25.5, 25.1, 22.4, 14.0; IR: 3133, 2931, 2097, 1458, 1362, 1196, 1046, 880; HRMS (ESI⁺, TOF) calcd for C₁₄H₂₇N₃NaO₂ [M + Na]⁺: 292.2001; found: 292.2003.

4-(4-(*tert*-Butylperoxy)butyl)-1-(4-phenylbutyl)-1H-1,2,3-triazole (29). By a similar procedure as for synthesis of **28** (above), reaction of a THF solution (3.5 mL) of 1-azido-4-(*tert*-butylperoxy)butane **12** (140 mg, 0.75 mmol) and 5-hexynyl benzene **15** (125 mg, 0.79 mmol), CuI (30.5 mg, 0.16 mmol, 0.2 equiv.), and triethylamine (0.32 mL, 2.30 mmol, 3 equiv.) furnished 38.5 mg (15%) of peroxy triazole **29** as a light yellow oil: R_f =



0.23 (40% EA/Hex); ^1H δ 7.31–7.17 (m, 6H), 4.39 (t, J = 7.3, 2H), 3.99 (t, J = 6.2, 2H), 2.77–2.64 (m, 4H), 2.02 (m, 2H), 1.83–1.61 (m, 6H), 1.26 (s, 9H); ^{13}C δ 152.1, 142.5, 128.6, 128.5, 128.4, 125.8, 80.3, 74.0, 50.6, 35.7, 31.0, 28.7, 27.0, 26.5, 26.1, 25.1; IR: 3026, 2925, 2155, 2032, 1454, 1362, 1196, 699; HRMS (ESI⁺, TOF) calcd for $\text{C}_{20}\text{H}_{31}\text{N}_3\text{NaO}_2$ [$\text{M} + \text{Na}$]⁺: 368.2314; found: 368.2313.

4-Butyl-1-(4-(*tert*-butylperoxy)butyl)-1*H*-1,2,3-triazole (28, CuSO_4 catalyst). To a solution of azidoperoxide **12** (0.80 mmol) and 1-hexyne **14** (0.9 mmol) in CH_2Cl_2 (3.5 mL) was added a solution of CuSO_4 (0.16 mmol, 0.2 equiv.) and sodium ascorbate (0.16 mmol, 0.2 equiv.) in water (1.5 mL). After stirring for 24 h, the reaction was diluted with water. The combined ethyl acetate extracts (3×15 mL) were dried with Na_2SO_4 , concentrated under reduced pressure and purified by column chromatography (10% EA/Hex) to yield 61.7 mg (29%) of peroxy triazole **28** as a light yellow oil: R_f = 0.25 (40% EA/Hex).

4-(4-(*tert*-Butylperoxy)butyl)-1-(4-phenylbutyl)-1*H*-1,2,3-triazole (29, CuSO_4 catalyst). By a similar procedure as described above for synthesis of **28**, reaction of azidoperoxybutane **12** and hexynyl benzene (**15**) with CuSO_4 and sodium ascorbate in CH_2Cl_2 water furnished 7.1 mg (5%) of peroxy triazole **29**.

Unsuccessful click reaction involving silylperoxy/azide **13**

Reaction of peroxide **13** with alkynes **14** or **15** in the presence of 20% $\text{CuI}/\text{Et}_3\text{N}$ under conditions similar to those described earlier furnished a mixture of peroxide decomposition products. The same reaction, when conducted in the presence of 20% CuSO_4 /sodium ascorbate resulted in no detectable reaction (TLC) and extensive recovery of **13**.

Preparation of functionalized nanoparticles

Synthesis of pentanethiolate functionalized gold nanoparticles was adapted from a reported procedure.^{38a} An aqueous solution of hydrogen tetrachloroaurate (10 mL, 30 mmol, 30 mM) was vigorously stirred with a toluene solution of tetraoctylammonium bromide (27 mL, 50 mmol, 50 mM) until the tetrachloroaurate was transferred into the organic toluene layer. Pentanethiol (42 mg) was then added to the organic phase, followed by the addition of freshly prepared aqueous solution of sodium borohydride (8.33 mL, 3.3 mmol, 0.4 M). After the reaction had stirred for 3 hours at room temperature, the organic layer was separated and concentrated under reduced pressure to a volume of ~ 1 mL. The concentrated layer was mixed with 125 mL of methanol and held at -18°C for 4 hours. The resulting solution was centrifuged. The nanoparticle-containing pellet was resuspended (methanol, 25 mL) and again centrifuged ($3\times$) to yield (73.1 mg) a brown precipitate: ^1H δ 1.32 (4H), 0.92 (3H);³⁹ IR: 2956, 2921; elemental analysis: Au, 50.1; C, 43.9; S, 6.0%; XPS (binding energies): Au $4f_{7/2}$ (82.6 eV), Au $4f_{5/2}$ (86.4 eV), C 1s (284.3 eV), S 2p (161.2 eV). Solutions of the nanoparticles displayed a broad absorption in the following paragraph, a dodecanethiol-passivated set of nanoparticles prepared as a standard, and which did display the small 520 nm bump, otherwise exhibited identical properties to the C5 nanoparticles.⁷¹

Dodecanethiolate functionalized gold nanoparticles.

Synthesis of dodecanethiolate functionalized gold nanoparticles were prepared following a reported procedure.^{38a} An aqueous solution of hydrogen tetrachloroaurate (10 mL, 30 mmol, 30 mM) was vigorously stirred with a toluene solution of tetraoctylammonium bromide (27 mL, 50 mmol, 50 mM) until the tetrachloroaurate was transferred into the organic toluene layer. Dodecanethiol (57 mg) was then added to the organic phase, followed by the addition of freshly prepared aqueous solution of sodium borohydride (8.33 mL, 3.3 mmol, 0.4 M). The reaction was stirred for 3 hours at room temperature. After 3 hours of stirring at room temperature, the organic layer was separate, and concentrated under reduced pressure to 1 mL, and mixed with 125 mL of methanol to remove excess thiol and kept at -18°C for 4 hours. The solution was centrifuged and washed with methanol ($25\text{ mL} \times 3$) to yield (73.1 mg) a black precipitate: ^1H δ 1.27 (10H), 0.91 (3H); IR: 2956, 2920; UV-vis: 520.65 nm (diagnostic signal; see ESI⁺); elemental analysis: Au, 78.4; C, 17.7; S, 3.9%; XPS (binding energies): Au $4f_{7/2}$ (82.9 eV), Au $4f_{5/2}$ (86.6 eV), C 1s (283.8 eV), S 2p (161.0 eV).

Azide functionalized gold nanoparticles (N_3Au) were prepared *via* a modification of reported procedures.³⁶ To a solution of $\text{C}_5\text{SH-Au P}$ (47.9 mg) in CH_2Cl_2 (5 mL) was added 6-azido-hexane-1-thiol (142.5 mg). After stirring for 3 days at room temperature under nitrogen, the reaction solution was concentrated under reduced pressure and pelleted/resuspended (6×10 mL, methanol) to yield (62.3 mg) a brown precipitate: ^1H δ 3.31 (2H), 1.45 (2H), 1.28 (4H), 0.87 (3H); IR: 2922, 2851, 2094; elemental analysis: Au, 54.0; C, 34.2; S, 8.5; N, 3.3%; XPS (binding energies): Au $4f_{7/2}$ (83.1 eV), Au $4f_{5/2}$ (86.9 eV), C 1s (283.9 eV), S 2p (162.0 eV), N 1s (398.3 eV).

Peroxide functionalized gold nanoparticles (**PeroxideAu**) were prepared *via* a modification of reported procedures.³⁷ To a solution of 6-(*tert*-butylperoxy)-1-hexyne (182.9 mg, 1.07 mmol) and azidoalkylthiolate Au-NP (23.4 mg) in THF (5 mL) was added CuI (41.7 mg, 0.22 mmol, 0.2 equiv. per alkyne), followed by triethylamine (0.50 mL, 3.6 mmol, 3 equiv. per alkyne). The reaction was stirred at room temperature for 8 hours. The reaction was diluted with saturated aqueous solution of NH_4Cl (10 mL) and the organic layer was washed with water ($10\text{ mL} \times 3$). The organic layer was dried with Na_2SO_4 , concentrated under reduced pressure and washed with methanol ($10\text{ mL} \times 6$) to remove excess alkyne to yield (29.8 mg) a light brown precipitate: ^1H δ 3.99 (2H), 2.47 (2H), 1.60–1.75 (8H), 1.26–1.30 (11H), 0.88 (1.5 H) IR: 2975, 2926, 2865, 1723, 1559, 1463, 1362, 1022, 803; elemental analysis: Au, 26.4; C, 49.9; S, 4.7; N, 5.1; O, 13.9%; XPS (binding energies): Au $4f_{7/2}$ (84.0 eV), Au $4f_{5/2}$ (87.8 eV), C 1s (284.3 eV), S 2p (162.9 eV), N 1s (399.6 eV), O 1s (531.5 eV).

Peroxyacetal functionalized gold nanoparticles (Perox-yacetalAu). To a solution of 6-((2-peroxypropan-2-yl)oxy)-1-hexyne (197.9 mg, 1.06 mmol) and azidoalkylthiolate Au-NP (28.2 mg) in THF (5 mL) was added CuI (41.4 mg, 0.22 mmol, 0.2 equiv. per alkyne), followed by triethylamine (0.50 mL, 3.6 mmol, 3 equiv. per alkyne). The reaction was stirred at room temperature for 8 hours. The reaction was diluted with



saturated aqueous solution of NH_4Cl (10 mL) and the organic layer was washed with water (10 mL $\times$ 3). The organic layer was dried with Na_2SO_4 , concentrated under reduced pressure and washed with methanol (10 mL $\times$ 6) to remove excess alkyne to yield (14.1 mg) a yellow precipitate: ^1H δ 3.86 (2H), 2.05 (2H), 1.67 (2H), 1.57 (6H), 1.39 (5 H), 1.28 (4H), 1.12 (4H), 0.91 (5H); IR: 3411, 2921, 2851, 1738, 1240, 1065; elemental analysis: Au, 20.0; C, 52.6; S, 4.3; N, 4.6; O, 18.5%; XPS (binding energies): Au 4f_{7/2} (83.9 eV), Au 4f_{5/2} (87.6 eV), C 1s (284.0 eV), S 2p (163.2 eV), N 1s (399.3 eV), O 1s (531.5 eV).

Abbreviations

Hexane(s)	Hex
EA	Ethyl acetate
THF	Tetrahydrofuran
DMF	<i>N,N</i> -Dimethylformamide
DMSO	Dimethyl sulfoxide
EtOH	Ethanol
MeOH	Methanol
TEA	Triethylamine
Na ascorb.	Sodium ascorbate
TBHP	<i>tert</i> -Butyl hydroperoxide
TBTA	Tris[(1-benzyl-1 <i>H</i> -1,2,3-triazol-4-yl)methyl]amine
Co(acac) ₂	Cobalt(II) acetylacetonate

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

Funding was provided by NSF (CHE 1464914) and by Dr In Quen Lee. Materials characterization was performed in part in the Nebraska Nanoscale Facility: National Nanotechnology Coordinated Infrastructure and the Nebraska Center for Materials and Nanoscience, facilities supported by NSF (ECCS 1542182) and the Nebraska Research Initiative. We thank Prof. Martha Morton, Ms Nataliia Vorobyova, Dr Moriah Locklear, and Mr Boone Evans for technical assistance and/or insightful suggestions.

References

- (a) I. A. Yaremenko, V. A. Vil, D. V. Demchuk and A. O. Terent'ev, Rearrangements of organic peroxides and related processes, *Beilstein J. Org. Chem.*, 2016, **12**, 1647–1748; (b) P. Dussault, Reactions of Hydroperoxides and Peroxides, in *Active Oxygen In Chemistry*, Chapman & Hall, London, 1995, pp. 141–203; (c) *The Chemistry of Peroxides*, ed. J. F. Liebman and A. Greer, Wiley, 2014, vol. 3, p. 1120, ISBN: 978-1-118-41271-8.
- For example, see: A. Banerjee and H. Yamamoto, Direct N–O bond formation via oxidation of amines with benzoyl peroxide, *Chem. Sci.*, 2019, **10**, 2124–2129.
- G. O. Wilson, J. W. Henderson, M. M. Caruso, B. J. Blaiszik, P. J. McIntire, N. R. Sottos, S. R. White and J. S. Moore, Evaluation of peroxide initiators for radical polymerization-based self-healing applications, *J. Polym. Sci., Part A: Polym. Chem.*, 2010, **48**, 2698–2708.
- P. M. O'Neill, V. E. Barton and S. A. Ward, The molecular mechanism of action of artemisinin—the debate continues, *Molecules*, 2010, **15**, 1705–1721.
- P. H. Dussault, A. D. George and T. K. Trullinger, Peroxides as oxidative enzyme inhibitors: mechanism-based inhibition of a cysteine protease by an amino acid ozonide, *Bioorg. Med. Chem. Lett.*, 1999, **9**, 3255–3258.
- S. Kyasa, R. N. Meier, R. A. Pardini, T. K. Truttmann, K. T. Kuwata and P. H. Dussault, Synthesis of Ethers via Reaction of Carbanions and Monoperoxyacetals, *J. Org. Chem.*, 2015, **80**, 12100–12114.
- (a) N. A. Porter, A perspective on free radical autoxidation: the physical organic chemistry of polyunsaturated fatty acid and sterol peroxidation, *J. Org. Chem.*, 2013, **78**, 3511–3524; (b) A. Bodin, M. Linnerborg, J. L. G. Nilsson and A.-T. Karlberg, Structure Elucidation, Synthesis, and Contact Allergenic Activity of a Major Hydroperoxide Formed at Autoxidation of the Ethoxylated Surfactant C12E5, *Chem. Res. Toxicol.*, 2003, **16**, 575–582.
- (a) T. M. McIntire, O. Ryder and B. Finlayson-Pitts, Secondary ozonide formation from the ozone oxidation of unsaturated self-assembled monolayers on zinc selenide attenuated total reflectance crystals, *J. Phys. Chem. C*, 2009, **113**, 11060–11065; (b) A. Razgon, R. G. Bergman and C. N. Sukenik, Ozonolysis-based route to the in situ formation of aldehyde-bearing self-assembled monolayer surfaces, *Langmuir*, 2008, **24**, 2545–2552.
- (a) H. X. Jin, H. H. Liu, Q. Zhang and Y. Wu, On the susceptibility of organic peroxy bonds to hydride reduction, *J. Org. Chem.*, 2005, **70**, 4240–4247; (b) I. Opsenica, D. Opsenica, K. S. Smith, W. K. Milhous and B. A. Šolaja, Chemical stability of the peroxide bond enables diversified synthesis of potent tetraoxane antimalarials, *J. Med. Chem.*, 2008, **51**, 2261–2266; (c) P. Dussault, The peroxide changes everything: new methodology for the synthesis of peroxide-containing natural products, *Synlett*, 1995, 997–1003; (d) H. O'Dowd, P. Ploypradith, S. Xie, T. A. Shapiro and G. H. Posner, Antimalarial artemisinin analogs. Synthesis via chemoselective C–C bond formation and preliminary biological evaluation, *Tetrahedron*, 1999, **55**, 3625–3636; (e) P. H. Dussault and C. T. Eary, Palladium-Mediated Carbon–Carbon Bond Forming Reactions as a New Method for the Synthesis of Peroxides and Hydroperoxides, *J. Am. Chem. Soc.*, 1998, **120**, 7133–7134.
- (a) J. Zabicky, Analytical and Safety Aspects of Organic Peroxides and Related Functional Groups, in *The Chemistry of the Peroxide Group*, ed. Z. Rappoport, John Wiley & Sons, Chichester, 2006, vol. 2, pt 2, pp. 597–773; ISBN: 978-0-470-86274-2; (b) D. E. Clark, Peroxides and Peroxide-forming Compounds, *J. Chem. Health Saf.*, 2001, **8**, 12–21; (c) G. D. Kozak, A. N. Tsvigunov and N. I. Akinin, Similarities



- and Differences in Explosion Hazards of Organic Peroxides and High Explosives: an Experimental Study, *Cent. Eur. J. Energ. Mater.*, 2011, **8**, 249–260; (d) P. H. Dussault, *Working with Organic Peroxides in the Academic Lab*, Department of Chemistry, University of Nebraska-Lincoln; Digital Commons@University of Nebraska Lincoln; DOI: 10.13014/K2GT5KCJ; accessed July 2020.
- 11 (a) C. S. McKay and M. G. Finn, Click chemistry in complex mixtures: bioorthogonal bioconjugation, *Chem. Biol.*, 2014, **21**, 1075–1101; (b) W. H. Binder and R. Sachsenhofer, 'Click' chemistry in polymer and materials science, *Macromol. Rapid Commun.*, 2007, **28**, 15–54; (c) J. E. Moses and A. D. Moorhouse, The growing applications of click chemistry, *Chem. Soc. Rev.*, 2007, **36**, 1249–1262; (d) H. C. Kolb and K. B. Sharpless, The growing impact of click chemistry on drug discovery, *Drug Discovery Today*, 2003, **8**, 1128–1137; (e) R. Chelmowski, D. Kafer, S. D. Köster, T. Klasen, T. Winkler, A. Terfort, N. Metzler-Nolte and C. Wöll, Postformation modification of SAMs: using click chemistry to functionalize organic surfaces, *Langmuir*, 2009, **25**, 11480–11485.
 - 12 (a) M. Meldal and C. W. Tornøe, Cu-catalyzed azide–alkyne cycloaddition, *Chem. Rev.*, 2008, **108**, 2952–3015; (b) J. E. Hein and V. V. Fokin, Copper-catalyzed azide-alkyne cycloaddition and beyond: new reactivity of copper acetylides, *Chem. Soc. Rev.*, 2011, **39**, 1302–1315; (c) L. Zhu, C. J. Brassard, X. Zhang, P. M. Guha and R. J. Clark, On the Mechanism of Copper(I)-Catalyzed Azide–Alkyne Cycloaddition, *Chem. Rev.*, 2016, **16**, 1501–1517.
 - 13 (a) S. Jana, S. Iram, J. Thomas, S. Liekens and W. Dehaen, Synthesis and anticancer activity of novel aza-artemisinin derivatives, *Bioorg. Med. Chem.*, 2017, **25**, 3671–3676; (b) H. M. Ismail, V. E. Barton, M. Panchana, S. Charoensutthivarakul, G. A. Biagini, S. A. Ward and P. M. O'Neill, A Click Chemistry-Based Proteomic Approach Reveals that 1,2,4-Trioxolane and Artemisinin Antimalarials Share a Common Protein Alkylation Profile, *Angew. Chem., Int. Ed.*, 2016, **55**, 6401–6405; (c) W. Li, Y. Zhou, G. Tang and Y. Xiao, Characterization of the Artemisinin Binding Site for Translationally Controlled Tumor Protein (TCTP) by Bioorthogonal Click Chemistry, *Bioconjugate Chem.*, 2016, **27**, 2828–2833; (d) L. H. Binh, N. T. T. Van, V. T. Kien, N. T. T. My, L. Van Chinh, N. T. Nga, H. X. Tien, D. T. Thao and T. K. Vu, Synthesis and in vitro cytotoxic evaluation of new triazole derivatives based on artemisinin via click chemistry, *Med. Chem. Res.*, 2016, **25**, 738–750.
 - 14 (a) Y. Chen, J. Q. Ren, X. G. Zhang, D. Y. Wu, A. G. Shen and J. M. Hu, Alkyne-Modulated Surface-Enhanced Raman Scattering-Palette for Optical Interference-Free and Multiplex Cellular Imaging, *Anal. Chem.*, 2016, **88**, 6115–6119; (b) R. Kozłowski, A. Ragupathi and R. B. Dyer, Characterizing the Surface Coverage of Protein-Gold Nanoparticle Bioconjugates, *Bioconjugate Chem.*, 2018, **29**, 2691–2700; (c) S. Liu and M. Han, Synthesis, Functionalization, and Bioconjugation of Monodisperse, Silica-Coated Gold Nanoparticles: Robust Bioprobes, *Adv. Funct. Mater.*, 2005, **15**, 961–967.
 - 15 X.-R. Jiang, P. Wang, C. L. Smith and B. T. Zhu, Synthesis of Novel Estrogen Receptor Antagonists Using Metal-Catalyzed Coupling Reactions and Characterization of Their Biological Activity, *J. Med. Chem.*, 2013, **56**, 2779–2790.
 - 16 P. Dussault and A. Sahli, 2-Methoxy-2-propyl hydroperoxide: a convenient reagent for the synthesis of hydroperoxides and peracids, *J. Org. Chem.*, 1992, **57**, 1009–1012.
 - 17 K. Rajender Reddy, M. Venkateshwar, C. Uma Maheswari and P. Santhosh Kumar, Mild and efficient oxy-iodination of alkynes and phenols with potassium iodide and tert-butyl hydroperoxide, *Tetrahedron Lett.*, 2010, **51**, 2170–2173.
 - 18 (a) Y. Yamamoto, E. Niki and Y. Kamiya, Ozonization of Organic Compounds. Relative Reactivity of Protic Solvents Toward Carbonyl Oxide, *Bull. Chem. Soc. Jpn.*, 1982, **55**, 2405–2406; (b) T. Fisher and P. H. Dussault, Alkene Ozonolysis, *Tetrahedron*, 2017, **73**, 4233–4258.
 - 19 P. H. Dussault and D. R. Davies, Synthesis of 1,2-dioxanes, 1,2,4-trioxanes, and 1,2,4-trioxepanes via cyclizations of unsaturated hydroperoxyacetals, *Tetrahedron Lett.*, 1996, **37**, 463–466.
 - 20 The pK_a of propargyl alcohol is 13.2. Trifluoroethanol ($pK_a \sim 12.4$) has been employed as a nonnucleophilic polar solvent for ozonolysis: K. J. McCullough, H. Tokuhara, A. Masuyama and M. Nojima, New approaches to the synthesis of spiro-peroxylactones, *Org. Biomol. Chem.*, 2003, **1**, 1522–1527.
 - 21 Y. Hamada, H. Tokuhara, A. Masuyama, M. Nojima, H. S. Kim, K. Ono, N. Ogura and Y. Wataya, Synthesis and notable antimalarial activity of acyclic peroxides, 1-(alkyldioxy)-1-(methyldioxy) cyclododecanes, *J. Med. Chem.*, 2002, **45**, 1374–1378.
 - 22 L. Yao, B. T. Smith and J. Aubé, Base-Promoted Reactions of Bridged Ketones and 1,3- and 1,4-Haloalkyl Azides: Competitive Alkylation vs. Azidation Reactions of Ketone Enolates, *J. Org. Chem.*, 2004, **69**, 1720–1722.
 - 23 (a) S. Isayama and T. Mukaiyama, A New Method for Preparation of Alcohols from Olefins with Molecular Oxygen and Phenylsilane by the use of Bis(acetylacetonato) cobalt(II), *Chem. Lett.*, 1989, **18**, 1071–1074; (b) T. Tokuyasu, S. Kunikawa, A. Masuyama and M. Nojima, Co(III)-alkyl complex- and Co(III)-alkylperoxo complex-catalyzed triethylsilylperoxidation of alkenes with molecular oxygen and triethylsilane, *Org. Lett.*, 2002, **4**, 3595–3598.
 - 24 (a) *Organic azides: syntheses and applications*, ed. S. Bräse and K. Banert, John Wiley & Sons, 2010; (b) T. M. Klapötke, T. Rotariu, B. Stiasny, J. Stierstorfer, S. Wiegmann and T. Zecheru, Azido (tert-butylperoxy) methyl Compounds—An Exceptional Class of Energetic Materials, *Eur. J. Org. Chem.*, 2016, 4382–4386.
 - 25 (a) F. Himo, T. Lovell, R. Hilgraf, V. V. Rostovtsev, L. Noodlesman, K. B. Sharpless and V. V. Fokin, Copper(I)-Catalyzed Synthesis of Azoles. DFT Study Predicts Unprecedented Reactivity and Intermediates, *J. Am. Chem. Soc.*, 2005, **127**, 210–216; (b) P. Wu and V. V. Fokin,



- Catalytic Azide-Alkyne Cycloaddition: Reactivity and Applications, *Aldrichimica Acta*, 2007, **40**, 7–17.
- 26 L. L. Smith and F. L. Hill, Detection of sterol hydroperoxides on thin-layer chromatoplates by means of the Würster dyes, *J. Chromatogr.*, 1972, **66**, 101–109.
 - 27 (a) D. N. Barsoum, N. Okashah, X. Zhang and L. Zhu, Mechanism of Copper(I)-Catalyzed 5-Iodo-1,2,3-triazole Formation from Azide and Terminal Alkyne, *J. Org. Chem.*, 2015, **80**, 9542–9551; (b) D. N. Barsoum, C. J. Brassard, J. H. A. Deeb, N. Okashah, K. Sreenath, J. T. Simmons and L. Zhu, Synthesis of 5-Iodo-1,2,3-triazoles from organic azides and terminal alkynes: Ligand Acceleration Effect, Substrate Scope, and Mechanistic Insights, *Synthesis*, 2013, **45**, 2372–2386; (c) J. E. Hein, J. C. Tripp, L. B. Krasnova, K. B. Sharpless and V. V. Fokin, Copper (I)-catalyzed cycloaddition of organic azides and 1-iodoalkynes, *Angew. Chem., Int. Ed.*, 2009, **48**, 8018–8021.
 - 28 P. S. Donnelly, S. D. Zanatta, S. C. Zammit, J. M. White and S. J. Williams, “Click” cycloaddition catalysts: copper(i) and copper(ii) tris(triazolylmethyl)amine complexes, *Chem. Commun.*, 2008, 2459.
 - 29 (a) B. T. Worrell, J. A. Malik and V. V. Fokin, Direct Evidence of a Dinuclear Copper Intermediate in Cu(I)-Catalyzed Azide-Alkyne Cycloadditions, *Science*, 2013, **80**(340), 457–460; (b) R. Berg and B. F. Straub, Advancements in the mechanistic understanding of the copper-catalyzed azide-alkyne cycloaddition, *Beilstein J. Org. Chem.*, 2013, **9**, 2715–2750.
 - 30 B. H. Lipshutz, W. Chrisman and K. Noson, Hydrosilylation of aldehydes and ketones catalyzed by $[\text{Ph}_3\text{P}(\text{CuH})]_6$, *J. Organomet. Chem.*, 2001, **624**, 367–371.
 - 31 N. Kornblum and H. E. DeLaMare, The base catalyzed decomposition of a di-alkyl peroxide, *J. Am. Chem. Soc.*, 1951, **73**, 880–881.
 - 32 (a) E. W. Elliott, A. L. Ginzburg, Z. C. Kennedy, Z. Feng and J. E. Hutchison, Single-Step Synthesis of Small, Azide-Functionalized Gold Nanoparticles: Versatile, Water-Dispersible Reagents for Click Chemistry, *Langmuir*, 2017, **33**, 5796–5802; (b) D. Baranov and E. N. Kadnikova, Synthesis and characterization of azidoalkyl-functionalized gold nanoparticles as scaffolds for “click”-chemistry derivatization, *J. Mater. Chem.*, 2011, **21**, 6152.
 - 33 Y. Zhou, S. Wang, K. Zhang and X. Jiang, Visual Detection of Copper(II) by Azide- and Alkyne-Functionalized Gold Nanoparticles Using Click Chemistry, *Angew. Chem., Int. Ed.*, 2008, **47**, 7454–7456.
 - 34 J. L. Brennan, N. S. Hatzakis, T. R. Tshikhudo, N. Dirvianskyte, V. Razumas, S. Patkar, J. Vind, A. Svendsen, R. J. M. Nolte, A. E. Rowan and M. Brust, Bionanoconjugation via Click Chemistry: The Creation of Functional Hybrids of Lipases and Gold Nanoparticles, *Bioconjugate Chem.*, 2006, **17**, 1373–1375.
 - 35 V. Poonthiyil, T. K. Lindhorst, V. B. Golovko and A. J. Fairbanks, Recent applications of click chemistry for the functionalization of gold nanoparticles and their conversion to glyco-gold nanoparticles, *Beilstein J. Org. Chem.*, 2018, **14**, 11–24.
 - 36 P. Zhao, M. Grillaud, L. Salmon, J. Ruiz and D. Astruc, Click Functionalization of Gold Nanoparticles Using the Very Efficient Catalyst Copper(I) (Hexabenzyl)tris(2-aminoethyl)-amine Bromide, *Adv. Synth. Catal.*, 2012, **354**, 1001–1011.
 - 37 E. Boisselier, L. Salmon, J. Ruiz and D. Astruc, How to very efficiently functionalize gold nanoparticles by “click” chemistry, *Chem. Commun.*, 2008, 5788–5790.
 - 38 (a) M. Brust, M. Walker, D. Bethell, D. J. Schiffrin and R. Whyman, Synthesis of Thiol-derivatised Gold Nanoparticles in a Two-phase Liquid-Liquid System, *Chem. Commun.*, 1994, 801–802; (b) M. Brust, J. Fink, D. J. Schiffrin, D. Bethella and C. Kielyb, Synthesis and Reactions of Functionalised Gold Nanoparticle, *Chem. Commun.*, 1995, 1655–1656.
 - 39 Due to limited freedom of rotation, hydrogens near the Au-S interface are often invisible to solution phase NMR. (a) C. Guo and J. L. Yarger, Characterizing gold nanoparticles by NMR spectroscopy, *Magn. Reson. Chem.*, 2018, **56**, 1074–1082; (b) L. E. Marbella and J. E. Millstone, NMR Techniques for Noble Metal Nanoparticles, *Chem. Mater.*, 2015, **27**, 2721–2739.
 - 40 Oxford Instruments, *EDS in the TEM Explained (12 pages)*, 2013, <https://nano.oxinst.com/product/eds-for-tem>.
 - 41 X. Zhang, P. Liu and L. Zhu, Structural Determinants of Alkyne Reactivity in Copper-Catalyzed Azide-Alkyne Cycloadditions, *Molecules*, 2016, **21**, 1–17.
 - 42 A. Horn and P. H. Dussault, Synthesis of α -cyano and α -sulfonyl cyclic ethers via intramolecular reactions of peroxides with sulfone- and nitrile-stabilized carbanions, *J. Org. Chem.*, 2019, **84**, 14611–14626.
 - 43 A. S. Olson, A. J. Jameson, S. K. Kyasa, B. W. Evans and P. H. Dussault, Reductive Cleavage of Organic Peroxides by Iron Salts and Thiols, *ACS Omega*, 2018, **3**, 14054–14063.
 - 44 X. Wang, Y. Dong, S. Wittlin, D. Creek, J. Chollet, S. A. Charman, J. Santo Tomas, C. Scheurer, C. Snyder and J. L. Vennerstrom, Spiro- and dispiro-1,2-dioxolanes: contribution of iron (II)-mediated one-electron vs. two-electron reduction to the activity of antimalarial peroxides, *J. Med. Chem.*, 2007, **50**, 5840–5847.
 - 45 D. A. Fleming, C. J. Thode and M. E. Williams, Triazole Cycloaddition as a General Route for Functionalization of Au Nanoparticles, *Chem. Mater.*, 2006, **18**, 2327–2334.
 - 46 W. J. Sommer and M. Weck, Facile functionalization of gold nanoparticles via microwave-assisted 1,3 dipolar cycloaddition, *Langmuir*, 2007, **23**, 11991–11995.
 - 47 (a) W. Limapichat and A. Basu, Reagentless functionalization of gold nanoparticles via a 3 + 2 Huisgen cycloaddition, *J. Colloid Interface Sci.*, 2008, **318**, 140–144; (b) C. J. Thode and M. E. Williams, Kinetics of 1,3-dipolar cycloaddition on the surfaces of Au nanoparticles, *J. Colloid Interface Sci.*, 2008, **320**, 346–352.
 - 48 R. Cohen, Y. Mazuz, M. Tikhonov and C. M. Sukenik, Carboxylic Acid Decorated Self-Assembled Monolayer Films: New Acid Synthesis Chemistry and Reaction Chemistry Including Bridged Diacyl Peroxide Preparation, *Langmuir*, 2015, **31**, 3049–3058.



- 49 R. Mello and M. E. González Núñez, Reactions of Peroxides on Solid Surfaces, in *PATAI'S Chemistry of Functional Groups*, ed. Z. Rappoport, DOI: 10.1002/9780470682531.pat0880.
- 50 K. L. Norrod and K. L. Rowlen, Ozone-induced oxidation of self-assembled decanethiol: contributing mechanism for "photooxidation"? *J. Am. Chem. Soc.*, 1998, **120**(11), 2656–2657.
- 51 (a) W. Liu, X. Liu, D. Knaebel, L. Luck and Y. Li, Synthesis and antibacterial evaluation of novel water-soluble organic peroxides, *Antimicrob. Agents Chemother.*, 1998, **42**, 911–915; (b) C. G. Burkhart and C. N. Burkhart, Antibacterial properties of soluble benzoyl peroxide, *Int. J. Dermatol.*, 2008, **47**, 301–302.
- 52 For an example of surface-specific peroxide chemistry involving an adsorbed peroxide, see: N. Walalawela and A. Greer, Heterogeneous photocatalytic deperoxidation with UV and visible light, *J. Phys. Org. Chem.*, 2018, **31**, e3807.
- 53 K. Tokumaru, Complications in photochemistry of organic peroxides, *Res. Chem. Intermed.*, 1996, **22**, 255–273.
- 54 (a) R. Willand-Charnley, B. W. Puffer and P. H. Dussault, Oxacycle synthesis via intramolecular reaction of carbanions and peroxides, *J. Am. Chem. Soc.*, 2014, **136**, 5821–5823; (b) L. Lemee, M.-J. Bourgeois and E. Montaudon, Decomposition induced by acetylenic peroxides, *Bull. Soc. Chim. Belg.*, 1996, **105**, 467–472.
- 55 M. T. Włodarczyk, S. A. Dragulska, O. Camacho-Vanegas, P. R. Dottino, A. A. Jarzęcki, J. A. Martignetti and A. J. Mieszawska, Platinum (II) complex-nuclear localization sequence peptide hybrid for overcoming platinum resistance in cancer therapy, *ACS Biomater. Sci. Eng.*, 2018, **4**, 463–467.
- 56 (a) J. N. Eildal, A. Bach, J. Dogan, F. Ye, M. Zhang, P. Jemth and K. Strømgaard, Rigidified Clicked Dimeric Ligands for Studying the Dynamics of the PDZ1-2 Supramodule of PSD-95, *ChemBioChem*, 2015, **16**, 64–69; (b) A. Krasinski, Z. Radić, R. Manetsch, J. Raushel, P. Taylor, K. B. Sharpless and H. C. Kolb, In situ selection of lead compounds by click chemistry: target-guided optimization of acetylcholinesterase inhibitors, *J. Am. Chem. Soc.*, 2005, **127**, 6686–6692.
- 57 P. H. Dussault, I. Q. Lee, H. J. Lee, R. J. Lee, Q. J. Niu, J. A. Schultz and U. R. Zope, Peroxycarbenium-mediated C–C bond formation: applications to the synthesis of hydroperoxides and peroxides, *J. Org. Chem.*, 2000, **65**, 8407–8414.
- 58 J. E. Baldwin, L. Bischoff, T. D. W. Claridge, F. A. Heupel, D. R. Spring and R. C. Whitehead, An approach to the manzamine alkaloids modelled on a biogenetic theory, *Tetrahedron*, 1997, **53**, 2271–2290.
- 59 R. J. Armstrong, W. Niwetmarin and V. K. Aggarwal, Synthesis of Functionalized Alkenes by a Transition-Metal-Free Zweifel Coupling, *Org. Lett.*, 2017, **19**, 2762–2765.
- 60 J. van der Louw, J. L. van der Baan, F. J. de Kanter, F. Bickelhaupt and G. W. Klumpp, Allylmetallation of 1-silylalkynes by 2-(bromozincmethyl)-2-alkenyl ethers followed by Pd (0)-catalyzed cyclization: a one-pot synthesis of 4-methylenecyclopentenones, *Tetrahedron*, 1992, **48**, 6087–6104.
- 61 B. R. Brutiu, W. A. Bubeneck, O. Cvetkovic, J. Li and N. Maulide, On the formation of seven-membered rings by arene-ynamide cyclization, *Monatsh. Chem.*, 2019, **150**, 3–10.
- 62 M. Schmittel and C. Vavilala, Kinetic isotope effects in the thermal C₂–C₆ cyclization of enyne-allenes: experimental evidence supports a stepwise mechanism, *J. Org. Chem.*, 2005, **70**, 4865–4868.
- 63 J. M. Smith, T. Qin, R. R. Merchant, J. T. Edwards, L. R. Malins, Z. Liu, G. Che, Z. Shen, S. A. Shaw, M. D. Eastgate and P. S. Baran, Decarboxylative Alkynylation, *Angew. Chem., Int. Ed.*, 2017, **56**, 11906–11910.
- 64 J. Maury, L. Feray, M. P. Bertrand, A. Kapat and P. Renaud, Unexpected conversion of alkyl azides to alkyl iodides and of aryl azides to N-tert-butyl anilines, *Tetrahedron*, 2012, **68**, 9606–9611.
- 65 S. M. Smith and J. M. Takacs, Amide-Directed Catalytic Asymmetric Hydroboration of Trisubstituted Alkenes, *J. Am. Chem. Soc.*, 2010, **132**, 1740–1741.
- 66 T. Suzuki, Y. Ota, Y. Kasuya, M. Mutsuga, Y. Kawamura, H. Tsumoto, H. Nakagawa, M. G. Finn and N. Miyata, An unexpected example of protein-templated click chemistry, *Angew. Chem., Int. Ed.*, 2010, **49**, 6817–6820.
- 67 K. Motesharei and D. C. Myles, Molecular Recognition on Functionalized Self-Assembled Monolayers of Alkanethiols on Gold, *J. Am. Chem. Soc.*, 1998, **120**, 7328–7336.
- 68 M. H. Keylor, J. E. Park, C. J. Wallentin and C. R. J. Stephenson, Photocatalytic initiation of thiol-ene reactions: synthesis of thiomorpholin-3-ones, *Tetrahedron*, 2014, **70**, 4264–4269.
- 69 C. Ligeour, A. Meyer, J. J. Vasseur and F. Morvan, Bis- and Tris-Alkyne Phosphoramidites for Multiple 5'-Labeling of Oligonucleotides by Click Chemistry, *Eur. J. Org. Chem.*, 2012, 1851–1856.
- 70 F. González De Rivera, I. Angurell, O. Rossell, M. Seco and J. Llorca, Organometallic surface functionalization of gold nanoparticles, *J. Organomet. Chem.*, 2012, **715**, 13–18.
- 71 W. Haiss, N. T. K. Thanh, J. Aveyard and D. G. Fernig, Determination of Size and Concentration of Gold Nanoparticles from UV-Vis Spectra, *Anal. Chem.*, 2007, **79**, 4215–4221.

