



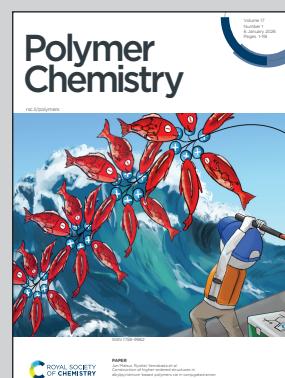
Showcasing research from Professor Dharmapura Murthy's laboratory, Department of Chemistry, Manipal Institute of Technology, Manipal Academy of Higher Education, Manipal, Karnataka, India.

A precursor-dependent distinctive polymerization process controls the optoelectronic properties of graphitic carbon nitride photocatalyst

This work provides mechanistic insights into how changing the precursor impacts the various intermediates formed during the thermal polymerization process of graphitic carbon nitride ( $g\text{-C}_3\text{N}_4$ ), a versatile photocatalyst. A comprehensive correlation is established between the polymerization mechanism and the optoelectronic, photophysical, structural, and photoelectrochemical properties of  $g\text{-C}_3\text{N}_4$ . Results aid in tailoring the properties of  $g\text{-C}_3\text{N}_4$  to enhance solar-to-chemical energy conversion efficiency by controlling the polymerisation process.

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# A precursor-dependent distinctive polymerization process controls the optoelectronic properties of graphitic carbon nitride photocatalyst

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Due to its metal-free polymeric nature, ease of synthesis using low-cost earth-abundant precursors and tunable optoelectronic properties, graphitic carbon nitride (GCN) is extensively used in solar fuel production. Despite two decades of extensive research, the fundamentals of the thermal polymerization process leading to the formation of GCN are inadequately understood. In this work, we employ cyanamide (CYN) and dicyandiamide (DCDA) precursors and systematically reveal the polymerization mechanism. Though CYN has half the amount of C and N compared with DCDA, it yielded virtually similar structural properties and a similar degree of conjugation that determines the energetic difference for  $\pi$ -to- $\pi^*$  fundamental (optical) transitions and photoexcited lifetimes. Detailed complementary analysis using thermal methods, along with quantifying the amount of  $\text{NH}_3$  released using the temperature-programmed desorption technique, offered unique insights into the polymerization process. Unlike previous notions, the results unambiguously demonstrate that GCN formation need not always release  $\text{NH}_3$  as a result of a thermal condensation reaction. Rather, it is possible that molecular rearrangement (dimerization and/or cyclization) of intermediate condensates can also play a major role in the formation of melamine, which is found to be an important intermediate. The obtained mechanistic insights into the thermodynamics of the polymerization process and its impact on optoelectronic properties and photoelectrochemical performance will aid the rational design of GCN to enhance the efficiency of solar energy conversion.

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## 1. Introduction

Among the wide range of semiconductor materials developed recently, graphitic carbon nitride (GCN) has gained significant attention due to its metal-free structure, ease of preparation, chemical stability, and absorption edge in the visible region, especially when compared to extensively used oxide-based photocatalysts.<sup>1</sup> In contrast to metal-based semiconductors that raise concerns related to toxicity, cost and long-term environmental impact, GCN offers a cost-effective alternative photocatalyst with tunable optoelectronic and surface properties. These prospects enable its application in photocatalytic  $\text{H}_2$  evolution,<sup>2,3</sup>  $\text{CO}_2$  reduction,<sup>4</sup> pollutant degradation<sup>5</sup> and organic transformation reactions.<sup>6,7</sup>

GCN can be easily synthesized *via* thermal condensation using a variety of cost-effective N-rich precursors like cyanamide (CYN),<sup>8</sup> dicyandiamide (DCDA),<sup>9</sup> melamine,<sup>10</sup> urea,<sup>11</sup> and thiourea.<sup>12</sup> Although GCN has been extensively employed for a wide range of applications in the fields of energy and

sustainability,<sup>13,14</sup> its basic polymerization mechanism has scarcely been investigated and remains unclear. Specifically, depending on the precursor and its C and N content, how the polymerization process changes and the nature of intermediates need to be studied in detail. It has been widely proposed that the thermal condensation of DCDA and CYN proceeds through a sequence of key intermediates like melamine, melam, melem and melon, which eventually form heptazine-based GCN.<sup>15</sup> However, probing how the formation of melamine, which is an important building block and functional unit of GCN, is impacted by precursor type and temperature is essential. Note that the degree of polymerization significantly affects defect formation in GCN and its optoelectronic properties, impacting a range of photoinduced processes such as charge photogeneration and transport and interfacial charge transfer. All these processes crucially impact the solar fuel production efficiency and can be tailored by the rational design of GCN; for this, a comprehensive understanding of the polymerization process plays a key role.

Similar to conducting polymers,  $\pi$  conjugation that is responsible for optical absorption and charge transport in GCN is offered by  $\text{sp}^2$ -hybridized C and N atoms in the heptazine ring. The highest occupied molecular orbital (HOMO) in

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**Fig. 1** (a) XRD patterns, (b) FTIR spectra, (c) optical absorption spectra measured in diffuse reflectance mode (inset: electronic transitions), and (d) time-resolved PL (tr-PL) transients of GCN-D and GCN-C. The dotted lines in the transient plots are a result of biexponential function fitting.

the N 1s core-level XPS spectra, which were deconvoluted into three peaks at 398.9, 400.3 and 401.4 eV, ascribed to C=N-C ( $N_1$  sites), N-(C) $_3$  ( $N_2$  sites) and -NH $_x$  ( $N_3$  sites), respectively.<sup>25</sup> The  $N_1/N_2$  peak area ratios from the deconvoluted N 1s spectra elucidate the sites of N-vacancies (either  $N_1$  or  $N_2$  sites). The  $N_1/N_2$  peak area ratios were found to be 3.01 and 5.41 for GCN-C and GCN-D, respectively. A decrease in the ratio proves that N-vacancies are present at  $N_1$  sites in heptazine units. Furthermore, Fig. 2d illustrates the EPR spectra of the synthesized samples. A sharp signal was observed at  $g = 2.004$ , which is attributed to unpaired electrons at sp $^2$ -hybridized C due to adjacent N-vacancies at two-coordinated nitrogen sites ( $N_1$  sites in Fig. 2c).<sup>26</sup> The higher signal intensity for GCN-C indicates a relatively higher concentration of N-vacancies compared to GCN-D. The N 1s XPS and EPR analyses collectively confirm the presence of a higher concentration of N-vacancies in GCN-C, which also results in the additional optical absorption tail (Fig. 1c).

Table 1 compares the results obtained for GCN-C and GCN-D from all analysis conducted. Despite changing the precursor from CYN to DCDA, which has half the amount of C and N, the bulk of the optoelectronic properties remain virtually similar. One of the most striking observations is the similar optical absorption onset at around 450 nm for both samples. This means that the degree of extended conjugation leads to an energetic difference between  $\pi$  and  $\pi^*$  orbitals of around 2.75 eV in both GCN-C and GCN-D. A minor yet significant 12% increase in the N-vacancy concentration for GCN-C is noticed compared to GCN-D. This observation agrees with the pronounced tail-state absorption of GCN-C between 450 and ~560 nm, which is attributed to electron transition from the HOMO to unoccupied N-vacancies below the LUMO (transition  $T_2$  in Fig. 1c). In short, notwithstanding the precursor and C/N content, most of the properties (structural and optoelectronic properties and excited-state behaviour) remain similar. In the following discussion, insights into such intriguing behaviour





**Fig. 2** Deconvoluted high-resolution C 1s core-level XPS spectra of (a) GCN-D and (b) GCN-C. (c) The structure of GCN highlighting C<sub>1</sub> and C<sub>2</sub> sites and (d) VB-XPS spectra of GCN-D and GCN-C.

will be detailed by invoking distinctive polymerization processes that can occur.

### 2.3. Unveiling precursor-dependent polymerization mechanisms to form GCN

Scheme 1 presents a widely adopted model for the formation of GCN *via* thermal condensation in a stepwise process involving multiple intermediates along with the release of NH<sub>3</sub> as a condensation byproduct.<sup>27,28</sup> Though the formation of GCN from melamine is straightforward, what remains unclear is how melamine is formed from CYN and DCDA precursors. This particular step is often overlooked and not investigated in detail. To answer this question, a detailed and complementary set of experiments was conducted. Differential scanning calorimetry (DSC) and thermogravimetric analysis (TGA) simultaneously offer information on the thermodynamics (exothermic or endothermic) of the polymerization process and on weight loss, respectively. Comparing this data will help to identify the polymerization process in different temperature regions. The unique aspect of this study is monitoring and quantifying the precursor-dependent release of NH<sub>3</sub> (as a condensation byproduct) using temperature-programmed desorp-

tion (NH<sub>3</sub>-TPD). Thus, the simultaneous correlation of TGA, DSC and NH<sub>3</sub>-TPD data at each temperature can aid in corroborating the polymerization process.

Let us first understand GCN formation from the DCDA precursor, which is now discussed. From DSC analysis of the precursor (Fig. 4a), the endothermic peak for DCDA at 207 °C corresponds to its melting, which is not accompanied by significant weight loss, as deduced from TGA (Fig. 4b). The prominent exothermic peak at 248–262 °C is ascribed to the cyclization of DCDA to form melamine. To confirm this assignment, when pristine melamine was studied using DSC, the exothermic peak at around 250 °C was absent (Fig. S3). Furthermore, the melting of melamine resulted in an endothermic peak at around 328–343 °C. Finally, the condensation of melamine into melem, which requires additional thermal energy, is evidenced by the endothermic peak at 394 °C. Previous studies predicted the release of NH<sub>3</sub> during the polymerization of DCDA,<sup>29</sup> and Fig. 4c shows significant release of NH<sub>3</sub> from DCDA at around 250 °C. As seen in TGA and DSC analysis, the peak at 250 °C is attributed to the formation of melamine. This similarity indicates that DCDA follows a widely adopted polymerization pathway to form me-





**Fig. 3** Deconvoluted high-resolution N 1s core-level XPS spectra of (a) GCN-D and (b) GCN-C. (c) The structure of GCN highlighting  $N_1$ ,  $N_2$  and  $N_3$  sites and (d) EPR spectra of GCN-D and GCN-C.

**Table 1** Summary of experimental observations from GCN-C and GCN-D samples

#### Similarities noticed

The bulk structural integrity revealed from XRD	GCN-D $\approx$ GCN-C
The fundamental $\pi$ -to- $\pi^*$ energy difference from the optical absorption onset	GCN-D $\approx$ GCN-C
The photoexcited state decay (lifetime) behaviour from tr-PL analysis	GCN-D $\approx$ GCN-C
The chemical functionality derived from FTIR	GCN-D $\approx$ GCN-C
The C constituents estimated from XPS	GCN-D $\approx$ GCN-C
The valence band onset from XPS	GCN-D $\approx$ GCN-C

#### Differences observed

The N-vacancy concentration deduced from XPS and EPR	GCN-C > GCN-D
The minor tail-state optical absorption from 450 to $\sim$ 560 nm	GCN-C > GCN-D

mine with the release of  $\text{NH}_3$ . Based on these experimental observations, a polymerization mechanism for DCDA is proposed in Fig. 4e. According to this scheme, two molecules of DCDA initially react by nucleophilic attack, further leading to cyclisation and subsequent resonance stabilization to form

melamine at around 248–262 °C. This step will eventually lead to the removal of carbodiimide, which decomposes to release  $\text{NH}_3$ , as confirmed by  $\text{NH}_3$ -TPD analysis (Fig. 4c).

GCN formation from the CYN precursor is proposed to occur as follows. The endothermic peak at 44 °C (absent for DCDA) corresponds to the melting of CYN, followed by an exothermic peak between 170 and 174 °C, which is attributed to the dimerization of CYN to form DCDA.<sup>30</sup> This latter process is also found to be accompanied by notable weight loss in the TGA scan (Fig. 4b). Upon the formation of DCDA, similar reactions are followed to form melamine. Comparing the thermal profiles of CYN and DCDA above 248 °C (Fig. 4a and b), similar GCN formation is expected from both precursors once melamine is formed. This observation suggests that melamine serves as the key intermediate and structural building block in the formation of GCN. Interestingly, unlike DCDA, the  $\text{NH}_3$ -TPD profile of CYN shows that a significantly lower amount of  $\text{NH}_3$  is released. This indicates that the polymerization pathway differs in the case of CYN up until the formation of melamine.

Based on this experimental evidence, we propose a detailed mechanism for the formation of melamine from CYN. The lone pair of electrons in the amine group of CYN attacks the electrophilic centre of another adjacent CYN molecule. Further





**Scheme 1** A generalised polymerization mechanism to form GCN involving an  $\text{NH}_3$  release model.

rearrangements and tautomerisation lead to the formation of DCDA (170–174 °C). The interaction of DCDA formed at 170–174 °C with other CYN molecules results in further cyclisation, leading to melamine formation. This set of reactions mainly involves rearrangements and thus does not produce significant byproducts, corroborating that a negligible amount of  $\text{NH}_3$  is released, as confirmed by  $\text{NH}_3$ -TPD analysis.

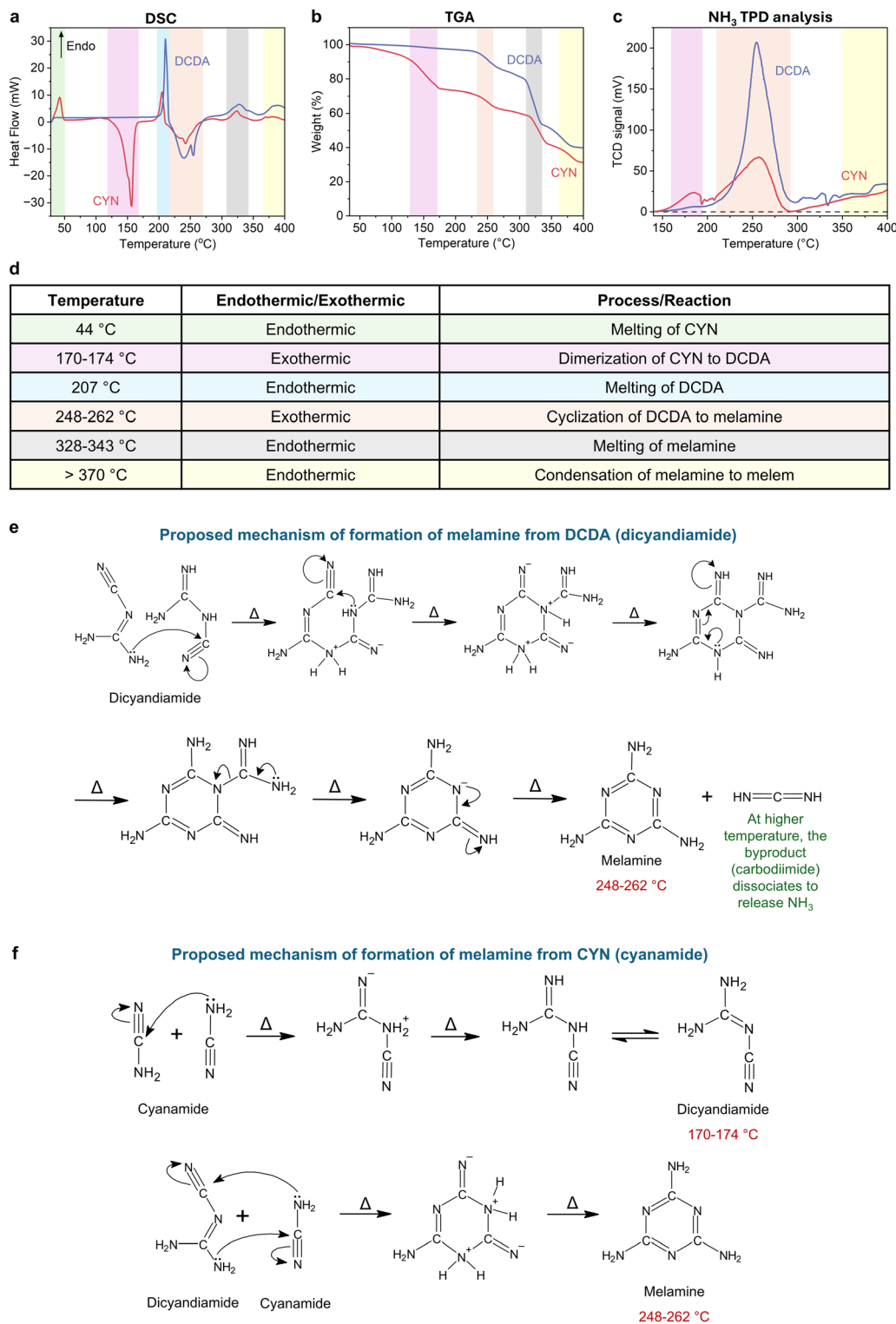
#### 2.4. Interplay between N-vacancies, extended $\pi$ conjugation, the polymerization process and PEC activity

As shown in Fig. 3, both XPS and EPR data confirmed that GCN-C has  $\sim 10\%$  more N-vacancies than GCN-D. With this information, we now aim to establish a plausible correlation between N-vacancy formation and the precursor-dependent polymerization process. Fig. 4b shows 25% weight loss for CYN during TGA in conjunction with a pronounced exothermic reaction, attributed to the dimerization of CYN to form DCDA. However, the proposed dimerization mechanism of CYN (Fig. 4f) is not expected to involve the release of  $\text{NH}_3$  or the loss of any C or N constituent atoms. This discrepancy suggests that some of the CYN precursor undergoes partial (self) decomposition to release  $\text{NH}_3$  (Fig. 4c, around 170 to 174 °C) and thereby forms intermediate condensates, potentially influencing the optoelectronic properties. The weak tail-state optical absorption due to  $T_2$  transitions (Fig. 1c) observed only for GCN-C further validates the role of incomplete condensates in affecting in-gap unoccupied N-vacancy formation. Though the experimental data strongly point towards our proposed model discussed earlier, further experiments are necessary to understand this intriguing phenomenon.

It is essential to understand the effect of the yield of GCN upon changing the precursor when commenting on the optoelectronic properties. The GCN yields obtained from CYN and DCDA precursors were found to be 42.6 and 50.8%, respectively. For both GCN-C and GCN-D, the fundamental  $\pi$ -to- $\pi^*$  transition at  $\sim 450$  nm is unaltered. This means that despite differences in the GCN yield and CYN having half the amount of C/N compared to DCDA, a minimum number of melamine units subsequently polymerized to form electronically connected functional heptazine units to endow the materials with a sufficient  $\pi$ -conjugation length. Thus, preserving the C and N content and not allowing the release of  $\text{NH}_3$  seems to be key for obtaining a virtually similar  $\pi$ -to- $\pi^*$  transition at  $\sim 450$  nm and functionality in GCN.

In our earlier discussion, the optoelectronic properties of GCN samples derived from both precursors were found to be virtually similar. To further evaluate these observations, the PEC performances of GCN-C and GCN-D are compared. Fig. 5a shows the CV data from GCN-D and GCN-C. Both samples exhibit similar CV profiles, with a higher current response on the cathodic side compared to the anodic side, demonstrating their potential for the  $\text{H}_2$  evolution reaction in agreement with earlier reports.<sup>31</sup> To confirm the cathodic behaviour, *i.e.*, that electrons are indeed the active species, CV experiments were conducted in the presence of  $\text{H}_2\text{O}_2$ , as depicted in Fig. 5b. The threefold enhancement in the cathodic current corroborates the notion that electrons are indeed involved in the probed electrochemical reaction. Fig. 5c and d show increment in photocathodic current by a factor of 1.23 and 1.09 for GCN-D and GCN-C samples, respectively. The electrochemically active



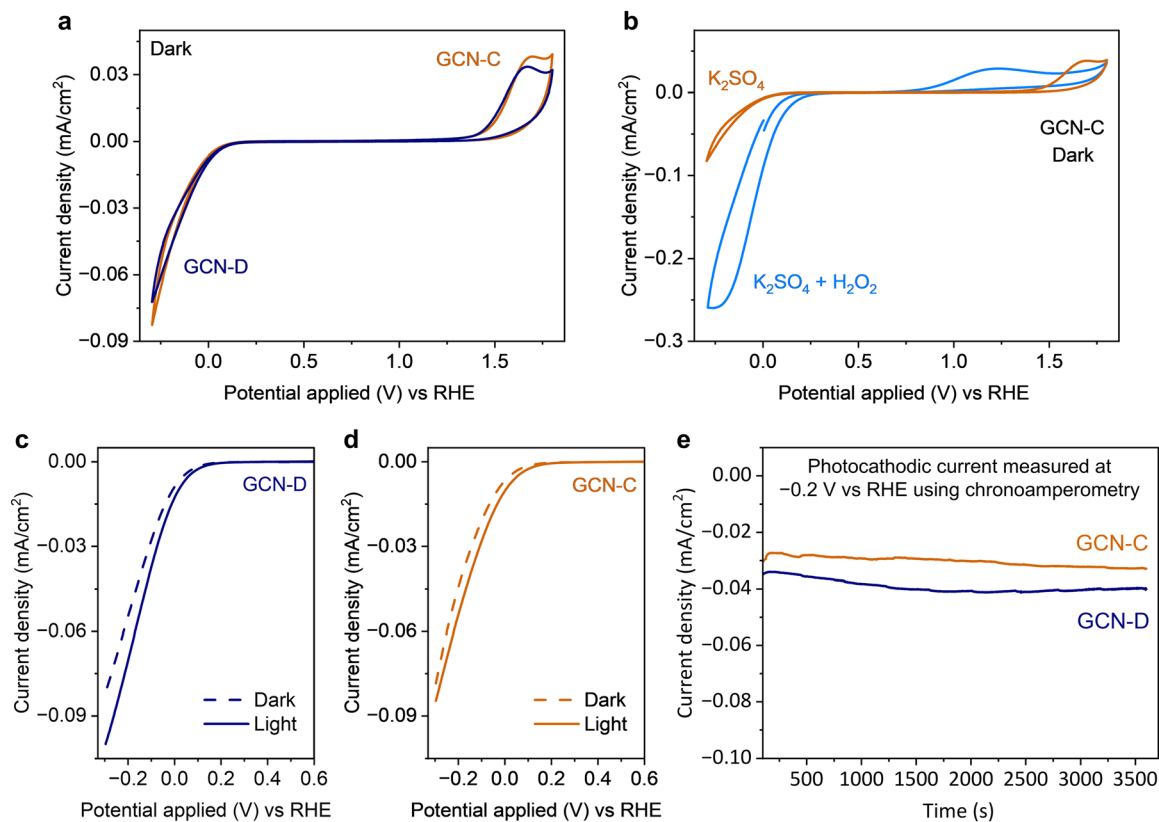


**Fig. 4** (a) DSC thermograms, (b) TGA thermograms and (c) NH<sub>3</sub>-TPD profiles of the DCDA and CYN precursors. (d) The processes/reactions involved in thermal polymerization to form GCN using DCDA and CYN. Schematic diagrams demonstrating the proposed formation of melamine from (e) DCDA and (f) CYN.

surface areas (ECSAs) of GCN-D and GCN-C are 1.20 and 1.29 cm<sup>2</sup> (Fig. S4), respectively. This suggests that both samples have similar ECSAs. However, the magnitude of the

photocathodic current of GCN-D is slightly higher compared to GCN-C, which may indirectly be attributed to an increased concentration of in-gap unoccupied N-vacancies that can





**Fig. 5** (a) Cyclic voltammetry (CV) data from GCN-D and GCN-C. (b) A comparison of CV data from GCN-C in the presence and absence of 0.3 mM H<sub>2</sub>O<sub>2</sub> as an electron acceptor. Linear sweep voltammetry (LSV) data from (c) GCN-D and (d) GCN-C under dark and light irradiation ( $\lambda > 420$  nm) conditions. All CV and LSV data are iR corrected. (e) Chronoamperometry measurements performed at  $-0.2$  V vs. RHE. Reaction conditions: AM 1.5 G; power density =  $1000$  W m<sup>-2</sup>; pH = 5.6; scan rate =  $50$  mV s<sup>-1</sup>.

potentially trap photogenerated electrons, leading to their inefficient migration towards the interface, thereby causing a marginal reduction in the photocathodic current of GCN-C compared to GCN-D. The chronoamperometry measurements shown in Fig. 5e indicate the prolonged stability of both these photoelectrodes measured up to one hour under light and bias. From a collective comparison of the data in Fig. 5, it is essential to note that both GCN-C and GCN-D, despite being formed from two different precursors *via* distinct polymerization mechanisms and notwithstanding the amount of NH<sub>3</sub> released, demonstrate comparable PEC performances.

Though the polymerization mechanism presented here is related to heptazine-based GCN, a similar model may well be applicable in other classes of carbon nitrides, such as poly (heptazine imides) (PHIs) and poly(triazine imides) (PTIs). These newly emerging families of 2D photocatalysts based on C and N have attracted significant attention for overall water splitting,<sup>32</sup> CO<sub>2</sub> reduction,<sup>33</sup> NH<sub>3</sub> production<sup>34</sup> and H<sub>2</sub>O<sub>2</sub> production.<sup>35</sup> These systems feature a delocalized  $\pi$ -electron system with triazine or heptazine units linked by imide groups and often require the use of eutectic salt melts with precursors to realize high performance.<sup>36</sup> However, it is unclear how polymerization happens in these systems, which significantly impacts the photocatalytic performance. The new approach

presented here may also be applied to elucidate correlations between polymerization mechanisms and photocatalytic activity. Such insights are essential when designing these materials to enhance the solar-to-chemical energy conversion efficiencies of these promising low-cost photocatalysts.

### 3. Conclusions

Despite the extensive utilization of GCN in photocatalytic reactions for sustainable applications and solar energy conversion, there is an inadequate understanding of how GCN is formed *via* thermal polymerization. Intrigued by this question, we systematically unveil the polymerization process using complementary thermal and NH<sub>3</sub>-TPD analyses for the first time and correlate our data with the observed optoelectronic/structural/surface properties. Such a comprehensive investigation demonstrated that the thermal polymerization process to form GCN need not always release NH<sub>3</sub> as a condensation byproduct. Overall, this work provides critical insights into the distinctive reaction pathways that form melamine building blocks *via* the polymerization of DCDA and CYN precursors. Correlating N-vacancy formation and the optoelectronic properties of GCN with the polymerization process and PEC



activity offered unique rational insights into tailoring the optoelectronic properties of GCN. These results may encourage further investigations, particularly those that have previously ignored the polymerization process, to gain a deeper understanding and thereby enhance the solar energy conversion efficiencies of these promising and cost-effective polymer-based photocatalysts.

## 4. Experimental section

### 4.1. Synthesis of GCN-D and GCN-C

Dicyandiamide (98%, Loba Chemie) and cyanamide (99%, Sigma-Aldrich) were used for synthesis without any further treatment. GCN was synthesized by heating DCDA at 550 °C for 3 h and is named GCN-D. GCN-C was also synthesized in a similar procedure using CYN as the precursor. To allow a comparison, the heating rate and duration were kept similar for both precursors.

### 4.2. Structural, surface and optoelectronic characterization tools

Powder XRD data were recorded using a Rigaku SmartLab diffractometer with Cu-K $\alpha$  radiation as the radiation source, with a wavelength ( $\lambda$ ) of 1.5406 Å. The sample was mounted on the sample holder, and the XRD pattern was recorded at a scan speed of 2.00° min<sup>-1</sup> with a 0.01° step width. FTIR spectroscopy was performed with a Shimadzu IRSpirit spectrophotometer in ATR mode. The optical absorption spectra of the synthesized samples were measured using a JASCO V-770 UV/Vis/NIR spectrometer equipped with an integrating sphere in diffuse reflectance mode. The steady-state PL measurements were carried out using an Edinburgh Instruments FLS1000 photoluminescence spectrometer. A 450-W ozone-free xenon arc lamp was used as the light source. For tr-PL measurements, the time-correlated single photon counting (TCSPC) technique was used. A picosecond pulsed diode laser of 405 nm was used as an excitation source, and a visible photomultiplier tube (PMT) was used as a detector. XPS analysis was carried out using a Thermo Scientific K-Alpha X-ray photoelectron spectrometer. Spectral charge correction was performed based on the C 1s peak appearing at 284.8 eV. EPR spectra were analyzed using a JEOL Model JES FA200 spectrometer. DSC measurements were conducted using a PerkinElmer DSC 6000 instrument. Equal amounts of precursor mixture were heated at a ramp rate of 5 °C min<sup>-1</sup> under argon atmosphere. TGA analysis was conducted using a PerkinElmer STA 6000 thermal analyzer. NH<sub>3</sub>-TPD analysis was carried out using the Autosorb 6200 EPDM XR instrument. The samples were heated at a ramp rate of 10 °C min<sup>-1</sup>, and the NH<sub>3</sub> released was quantified. The amounts of sample taken for analysis were as follows: DCDA = 0.020 g, CYN = 0.220 g.

### 4.3. Photoelectrochemical measurements

The working electrodes for PEC measurements were prepared as follows. 100 mg of photocatalyst (GCN-D and GCN-C) and 20 mg of ethyl cellulose were ground for 10 min. This mixture

was then added to a solution containing 1 mL of ethanol and 20  $\mu$ L of Nafion. The solution was sonicated for 2 h, followed by stirring it overnight to prepare a photocatalyst ink. 50  $\mu$ L of the prepared ink was spin-coated on ITO (indium tin oxide)-coated glass substrates. The electrodes were then calcined at 300 °C for 2 h. The PEC measurements were carried out using a potentiostat (Metrohm Autolab PGSTAT 204) and a solar simulator (Holmarc HR-SS300WRM1-100A) equipped with an air mass (AM 1.5 G) filter, and a 300-W xenon short arc lamp with a 420 nm long pass filter was used as the light source. A three-electrode system, with the ITO-coated glass substrate as the working electrode and a platinum and reversible hydrogen electrode (RHE) as the counter and reference electrodes, respectively, was used for the measurements. 0.1 M K<sub>2</sub>SO<sub>4</sub> (pH 5.6) was used as the electrolyte. CV scans were recorded over a potential range of -0.3 V to 1.8 V vs. RHE at a scan rate of 50 mV s<sup>-1</sup>. LSV analysis was performed from 0.8 V to -0.3 V vs. RHE at a scan rate of 50 mV s<sup>-1</sup>. The reported CV and LSV data involve iR-corrected potential values. The stability of the electrodes was studied using chronoamperometry measurements at -0.2 V vs. RHE for 1 hour. The ECSA was calculated by recording CV scans in the non-faradaic region at scan rates ranging from 10 mV s<sup>-1</sup> to 50 mV s<sup>-1</sup>. The double-layer capacitance ( $C_{dl}$ ) was obtained from a plot of anodic current ( $j_a$ )-cathodic current ( $j_c$ ) vs. the scan rate at a given potential. The ECSA was calculated using  $C_{dl}/C_s$ , where  $C_s$  is the specific capacitance of GCN.

## Author contributions

S. J. G. conducted the synthesis, structural/surface/optoelectronic characterization of the materials and analysis. S. J. G. and D. H. K. M. wrote the first version of the article. S. S. M. conducted thermal analysis measurements of the materials. S. S. M. and S. R. K. conducted the PEC measurements. S. R. K. deduced the polymerization mechanism. D. H. K. M. conceptualized the work, interpreted the results, edited the article, and coordinated the overall work. The article was written through contributions from all authors.

## Conflicts of interest

The authors declare no conflict of interest.

## Data availability

The data supporting this article have been included as part of the manuscript and supplementary information (SI). Supplementary information: XRD peak positions and calculated  $d$ -spacing values, steady-state PL emission spectra, PL lifetimes obtained from fitting, survey spectra from XPS, XPS peak parameters from the deconvolution of C and N 1s core-level spectra, and the DSC thermogram of melamine. See DOI: <https://doi.org/10.1039/d5py01003a>.



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