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Heterogeneous catalysis for oxygen tolerant photoredox atom transfer radical polymerization and small-molecule dehalogenation†

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Heterogeneous photocatalysts (PCs) have garnered attention for their sustainability and cost-effectiveness. Despite the existence of various types of these PCs, their synthesis often involves complex, multi-step procedures and laborious purification. Herein, we propose a simple method for attaching small-molecule photocatalytic species onto crosslinked 3-D polymer networks as insoluble scaffolds to create robust heterogeneous PCs. The highly swellable poly(ethylene glycol)-based ChemMatrix (CM) resin, known for its amphiphilic properties and high functional group loading, facilitated the covalent immobilization of the photoredox dye Eosin Y (EY), but also streamlined functionalization with Ir(III) complexes. The resulting heterogeneous CM-EY demonstrated efficient photocatalytic performance in open-to-air dual photoredox catalysis of atom transfer radical polymerization (photo-ATRP) under green light. This was confirmed by the well-controlled synthesis of polymers with molecular masses ranging from 20 kDa to 300 kDa and low dispersities. Furthermore, CM-EY exhibited excellent photostability and recyclability over multiple cycles of ATRP. The heterogeneous catalysis of photo-ATRP provided high temporal control and enabled benign conditions for synthesizing protein-polymer hybrids (PPH). When combined with the initiator-modified CM (CM-BIB), CM-EY facilitated the solid-phase synthesis of homopolymers and block copolymers with recyclable performance. However, the coordinatively bound Ir@CM showed decreased catalytic activity and efficiency toward photoredox dehalogenation due to the leaching of active species during recycling. This study highlights the advantages of the covalent linking of catalysts to solid supports over non-covalent interactions, underscoring the potential of functionalized polymer resin as a promising scaffold. Such an approach offers customization and tunability, presenting opportunities for innovation in green chemistry.

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Introduction

Photoredox catalysis (PCs) has emerged as a powerful tool in synthetic chemistry, allowing the activation of organic molecules under mild conditions through visible light absorption. This approach has changed organic transformations^{1–3} and facilitated reversible deactivation radical polymerizations (RDRPs) for the synthesis of macromolecules and polymer

bioconjugates.^{4–14} While homogeneous PCs offer advantages like accessible catalytic sites and high selectivity,^{14–18} recyclable heterogeneous PCs have gained prominence due to their versatility across various reaction mediums and cost-effectiveness.^{19–22} These recyclable PCs can undergo multiple catalytic cycles without compromising activity or selectivity, aligning well with green chemistry principles.^{23–25} Overall, the rational design of heterogeneous PCs represents a significant advancement, offering sustainable solutions and enhancing efficiency in various chemical processes.¹³

Photo-induced atom transfer radical polymerization (ATRP) has emerged as a commonly used technique in recent years due to its mild reaction conditions, ability to control polymerization temporally and spatially and oxygen tolerance.^{10,26–37} ATRP, a reversible redox process typically catalyzed by Cu complexes, involves the transfer of a halogen atom from the

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Polymer networks with 3D crosslinked architectures represent a vital class of functional soft materials.^{81–84} Gels act as swollen polymer networks, bridging the gap between solids and liquids.^{20,24} They have shown potential as heterogeneous catalyst supports due to their favorable swelling behaviors and flexibility.^{20,24} ChemMatrix (CM), a poly(ethylene glycol) (PEG)-based network initially developed for the solid-phase synthesis of peptides^{85,86} is comprised of thermally and chemically

In our study, we developed two heterogeneous photoactive materials based on CM-resin. Firstly, we covalently immobilized EY onto CM (CM-EY) for dual photoredox catalysis of ATRP. CM-EY demonstrated efficient photocatalytic performance, temporal control, and excellent photostability, ensuring prolonged catalytic activity over multiple cycles without degradation. Secondly, we anchored Ir(III) complexes onto ligand-functionalized CM (Ir@CM) through coordinative bonding. Ir@CM exhibited reduced efficiency and diminishing activity over successive cycles due to the leaching of active species. This study highlights the potential of functionalized polymer resin as a promising scaffold for catalyst immobilization due to its ease of separation. The importance of the covalent attachment of photocatalysts to solid supports and design principles and synthetic methodologies of these heterogeneous PCs are also discussed.

Eosin-functionalized ChemMatrix (CM-AM-EY)

Following the immobilization process, a fluorescence intensity counts (FIC) assay was conducted to quantify the concentration of EY immobilized per gram of both CM-AM-EY and CM-HMPB-EY (see ESI Fig. S1–S4†). The results indicated the presence of 4.6 ± 1.1 μmol EY per g of CM-AM-EY and $0.32 \pm$



Scheme 1 Immobilization of EY on CM solid support. (A) Using amine ($-\text{NH}_2$) functionalized sites ($0.5\text{--}0.7\text{ mmol g}^{-1}$) of CM resin (CM-AM). (B) Using $-\text{OH}$ functionalized sites ($0.4\text{--}0.6\text{ mmol g}^{-1}$) CM resin (CM-HMPB). (C) Digital images of CM-AM-EY and CM-HMPB-EY under ambient conditions and green light irradiation.

$0.06\text{ }\mu\text{mol EY per g}$ of CM-HMPB-EY. The digital images of the functionalized resins swollen in aqueous buffer under ambient conditions and green light irradiation revealed diminished fluorescence in the CM-HMPB-EY in contrast to CM-AM-EY (Scheme 1C). Despite using the same amount of EY for immobilization in both cases, the fluorescence intensity was 15 times lower for CM-HMPB-EY. This observation was counterintuitive as the CM resin shows better swelling performance in DMF, so better immobilization efficiency was expected. We hypothesized self-quenching of fluorescence in the CM-HMPB-EY due to the proximity of multiple EY molecules. Self-quenching of fluorescent dyes may occur due to non-covalent interactions such as π - π stacking or hydrophobic interactions. Furthermore, the Förster resonance energy transfer (FRET) mechanism could also lead to the non-radiative deactivation of the excited state of the dye molecules, resulting in reduced fluorescence intensity. To confirm our hypothesis, we cleaved the EY from CM-HMPB-EY support using 20% trifluoroacetic acid (TFA) (see ESI, Scheme S1†). The actual concentration of immobilized dye was $150 \pm 20\text{ }\mu\text{mol EY per g CM}$. Using these estimations, we computed the distance between EY molecules in CM-AM-EY and CM-HMPB-EY. The calculations revealed the average distance between EY molecules in CM-AM-EY to be 15.3 nm and 4.8 nm in CM-HMPB-EY (see ESI†). Since Förster resonance energy transfer (FRET) can typically occur over distances up to 10 nm, these calculations suggest EY may be quenched in CM-HMPB-EY by FRET. It should be noted that this analysis assumes that EY is ran-

domly distributed on the resin and does not aggregate. A more thorough characterization shall be needed to confirm the quenching phenomenon.

Optimization of heterogenous CM-AM-EY catalyzed photo ATRP

To verify the efficiency of CM-AM-EY to enable fully-oxygen tolerant dual photoredox/copper catalysis, oligo(ethylene glycol) methyl ether methacrylate (average $M_n = 500$, OEOMA₅₀₀) monomer was polymerized under green light (525 nm , 25.0 mW cm^{-2}) using 2-hydroxyethyl α -bromoisobutyrate (HOBIB) as the initiator, $\text{CuBr}_2/\text{TPMA}$ (TPMA = tris(2-pyridyl methyl)amine) as the catalyst and CM-AM-EY resin as the heterogeneous photoredox catalyst. The polymerizations were conducted in $1\times$ phosphate-buffered saline (PBS) with DMSO (10% v/v) in uncapped solid phase extraction (SPE) syringes equipped with $0.2\text{ }\mu\text{m}$ frits placed in a photo reactor with green LED lamp (525 nm , 25.0 mW cm^{-2}) at an ambient temperature maintained by a cooling fan. The polymerization reaction mixture was added to pre-weighed CM-AM-EY in an SPE cartridge and stirred at 500 rpm for 30 minutes, followed by irradiation for 30 minutes. After polymerization, the outlet of the syringe was uncapped, and the colorless, uncontaminated polymer sample was collected by simple vacuum filtration (Table 1).

In initial attempts, we performed heterogenous catalysis of photo ATRP using the previously optimized molar ratios $[\text{OEOMA}_{500}]/[\text{HOBIB}]/[\text{CM-AM-EY}]/[\text{CuBr}_2]/[\text{TPMA}] = 200/1/$



Table 1 Optimization of heterogeneous CM-EY catalyzed photo ATRP

Entry	CM-catalyst	[CM-EY] μM	[TEOA] mM	α_M^a (%)	$M_{n,th}^b$	$M_{n,app}^c$	$M_{n,abs}^d$	$\bar{D}$
1	CM-AM-EY	15	—	22%	22 000	12 500	16 800	1.16
2	CM-AM-EY	30	—	44%	44 000	22 500	32 400	1.13
3	CM-AM-EY	7.5	6.0	99%	99 000	60 900	94 500	1.15
4	CM-AM-EY	7.5	0.9	86%	86 000	58 500	88 200	1.17
5	CM-HMPB-EY	7.5	0.9	9%	9000	7600	9200	1.11
6	CM-HMPB-EY	7.5	3.0	88%	88 000	50 200	70 100	1.22

Reactions conditions: [OEOMA]/[HOBIB]/[CuBr₂]/[TPMA]/[CM-AM-EY]/[TEOA]: 200/1/0.2/0.6/XX/XX, [OEOMA₅₀₀] = 300 mM, [HOBIB] = 1.5 mM, in 1× PBS buffer, irradiated for 30 min under green light LEDs (λ_{max} = 520 nm, 25.0 mW cm⁻²), in an ambient atmosphere. Reaction volume 4.5 mL stirring at 500 rpm. ^a OEOMA₅₀₀ conversion was determined by using ¹H NMR spectroscopy. ^b Theoretical molecular weight was calculated using the equation $M_{n,th} = [M] \times MW_M \times \alpha_M + MW_I$. ^c Apparent molecular weight ($M_{n,app}$) was analyzed by SEC using DMF as the eluent and PMMA as calibration standards. ^d $M_{n,abs}$ analyzed using SEC run using 1× PBS as eluent equipped with triple detectors: multi-angle light scattering (MALS), refractive index (RI), and UV detector.

0.01/0.2/0.6) (Table 1, entry 1). However, the monomer conversion after 30 minutes of irradiation, as analyzed by ¹H NMR, was considerably lower (22%) compared to the previously reported 84% conversion of monomer under identical conditions using homogeneous EY as the photoredox catalyst.¹⁸ To enhance the polymerization rate, we doubled the molar ratio of CM-AM-EY (Table 1, entry 2). Although this adjustment led to an improvement in the overall monomer conversion after 30 minutes (44%), it remained relatively low, resulting in a significant amount of unreacted monomer in the final product.

The proposed mechanism from dual photoredox/copper catalysis suggested the oxidative quenching of excited state EY as the preferred pathway. Additionally, it highlighted the critical role of excess ligands in reducing the oxidized EY and closing the photocatalytic cycle.¹⁴ We envisioned that the addition of excess triethanolamine (TEOA), a more cost-effective alkylamine, as a sacrificial electron donor would accelerate the overall polymerization kinetics. Heterogeneous CM-AM-EY catalyzed photo ATRP was then performed in the presence of 3 mM of triethanolamine (TEOA), resulting in quantitative monomer conversion (>99%) and the SEC analysis revealed the absolute molecular weight ($M_{n,abs}$) of the polymer was slightly lower than the theoretical value ($M_{n,th}$) (Table 1, entry 3). Next, to enhance control, the polymerization was conducted with reduced amounts of CM-AM-EY and TEOA using molar ratios ([OEOMA₅₀₀]/[HOBIB]/[CM-AM-EY]/[CuBr₂]/[TPMA]/[TEOA]) = 200/1/0.005/0.2/0.6/0.6 still resulted in a high monomer conversion and the synthesized POEOMA₅₀₀ showed good agreement between theoretical and absolute molecular weight and a narrow molecular weight distribution (Table 1, entry 4). Therefore, these optimized conditions using CM-AM-EY were used in all the subsequent polymerizations.

CM-HMPB-EY was also tested for heterogeneous catalysis of photo ATRP and revealed lower catalytic efficiency than its counterpart CM-AM-EY (Table 1, entry 5). Using an excess of amines (TEOA) accelerated the rate of polymerization reaching quantitative monomer conversion. However, poor agreement between $M_{n,th}$ and $M_{n,abs}$ was observed (Table 1, entry 6).

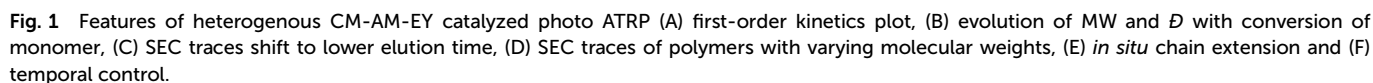
Kinetics of heterogeneous CM-AM-EY catalyzed photo ATRP

The kinetic analysis of heterogeneous CM-AM-EY catalyzed photo ATRP under optimized conditions ([OEOMA₅₀₀]/[HOBIB]/[CM-AM-EY]/[CuBr₂]/[TPMA]/[TEOA]) = 200/1/0.005/0.2/0.6/0.6 revealed a linear first-order relationship between $\ln([M]_0/[M]_t)$ and time. A short induction period of 5 min was followed by a rapid polymerization that reached 90% monomer conversion within 40 min (Fig. 1A). In addition, the molecular weights increased as a function of monomer conversion (Fig. 1B), and molecular weight distribution values remained low ($\bar{D} \leq 1.17$) during the polymerization (Fig. 1C).

Varying the targeted degree of polymerization (DP)

Polymers with varying molecular weights were synthesized by adjusting the degree of polymerization (DP) through systematic variation of HOBIB initiator concentration (6 mM to 0.30 mM), while maintaining constant polymerization conditions ([OEOMA₅₀₀] = 300 mM, [CM-AM-EY] = 7.5 μM , [CuBr₂] = 0.3 mM, [TPMA] = 0.9 mM, [TEOA] = 0.9 mM). For a broad DP range (50–800), monomer conversions reached $\geq 70\%$ within 30 minutes, yielding polymers with narrow molecular weight distributions ($\bar{D} = 1.08$ – 1.28) (Table 2, entries 1–5). At DP = 1000, conversion slightly decreased (66%), accompanied by a minor loss in molecular weight control (Table 2, entry 6). SEC traces exhibited symmetry without tailing or high molecular





Entry	[M]/[I]/[CM-AM-EY]/[CuBr ₂]/[TPMA]/[TEOA]	[I] (mM)	Conv. ^a (%)	<i>M</i> _{n,th} ^b	<i>M</i> _{n,abs} ^c	<i>D</i>
1	50/1/0.0025/0.05/0.15/0.15	6.0	89	22 250	27 800	1.08
2	100/1/0.005/0.1/0.3/0.3	3.0	88	44 000	52 400	1.05
3	200/1/0.01/0.2/0.6/0.6	1.5	86	86 000	92 200	1.05
4	400/1/0.02/0.4/1.2/1.2	0.75	82	164 000	172 800	1.08
5	800/1/0.04/0.8/2.4/2.4	0.375	72	288 000	235 000	1.28
6	1000/1/0.05/1.0/3.0/3.0	0.30	66	330 000	280 400	1.42

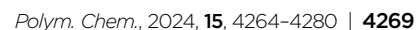
Reactions conditions: $[M]/[HOBIB]/[CuBr_2]/[TPMA]/[CM-AM-EY]/[TEOA]$: 200/1/0.2/0.6/0.005/0.6, $[M] = [OEGMA_{500}] = 300$ mM, $[HOBIB] = 6\text{--}0.30$ mM, $[TEOA] = 0.9$ mM in $1\times$ PBS buffer at rt, irradiated for 30 min under green light LEDs ($\lambda_{\text{max}} = 520$ nm, 25.0 mW cm^{-2}), stirring at 500 rpm in an ambient atmosphere. ^a $OEGMA_{500}$ conversion was determined by using ^1H NMR spectroscopy. ^b Theoretical molecular weight was calculated using the equation $M_{n,\text{th}} = [M] \times MW_M + \alpha M + MW_I$. ^c $M_{n,\text{abs}}$ analyzed using SEC run using $1\times$ PBS as eluent equipped with triple detectors: multiangle light scattering (MALS), refractive index (RI), and UV detector.

The chain-end fidelity was confirmed by performing an *in situ* chain extension experiment. The first block of POEOMA₅₀₀ was synthesized with molar ratios [OEOMA₅₀₀]/[HOBIB]/[CM-AM-EY]/[CuBr₂]/[TPMA]/[TEOA]: 50/1/0.0025/0.05/0.15/0.15 (conv. = 64%, $M_{n,abs}$ = 15 500, D = 1.26). POEOMA₅₀₀ was then used as a macroinitiator (macro-I) to graft the second block with target DP = 250. The polymerization was irradiated for 30 minutes resulting in a chain-extended POEOMA₅₀₀ (conv. = 74%, $M_{n,abs}$ = 115 400, D = 1.10). The SEC analysis

To demonstrate temporal control in the heterogeneous CM-AM-EY catalyzed photo ATRP, the polymerization mixture was subjected to intermittent light exposure with continued stirring at 500 rpm. The heterogeneous catalytic system showed high temporal control (Fig. 1F). Sustained polymerization was observed upon irradiation under green light while the dark periods halted the polymerization instantly and effec-

Therefore, we functionalized two different proteins, namely bovine serum albumin (BSA) and chymotrypsin (CT) with ATRP initiators to prepare BSA-iBBR₂₂ and CT-iBBR₁₂ respectively. BSA was chosen due to its relatively higher resilience while CT is a therapeutic enzyme used to treat pancreatic insufficiency. Grafting-from ATRP was used to graft PCBMA polymers from the protein's surface using heterogeneous CM-EY as the photocatalyst. The SEC-MALS analysis of

The monomer scope was expanded to include hydrophilic methacrylates known for their biocompatibility, such as carboxy betaine methacrylate (CBMA), (methylsulfanyl)ethyl methacrylate (MSEMA), and methacrylate with a shorter poly (ethylene glycol) unit (OEOMA₃₀₀). The resulting polymers – POEOMA₃₀₀, PCBMA, and PMSEMA, synthesized *via* heterogeneous CM-EY photo-ATRP exhibited predictable molecular weights and low dispersity (Table S1,[†] entries 1–3, Fig. 3A–C).



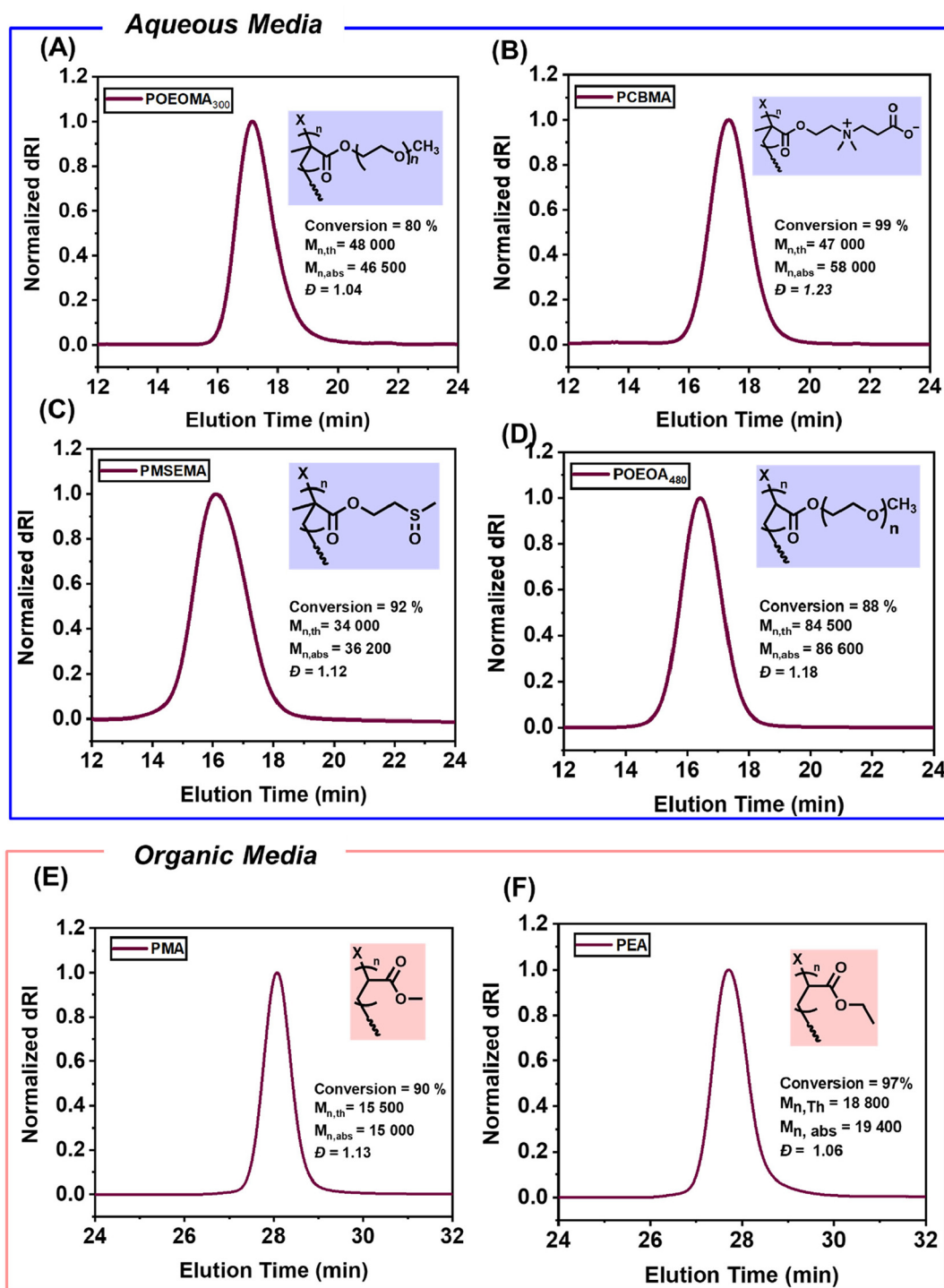


Fig. 3 Expanding monomer scope of heterogeneous CM-EY catalyzed photo ATRP to (A) POEOMA₃₀₀, (B) PCBMA, (C) PMSEMA, (D) POEOA₄₈₀, (E) PMA and (F) PEA.

the synthesized bioconjugates showed monomodal distribution and higher absolute molecular weights of the bioconjugates as compared to native proteins, indicating a successful grafting-from ATRP (Fig. 4A and B). The larger dispersities could be related to variable number of initiators anchored to proteins.

Solid-phase polymer synthesis by heterogeneous CM-EY catalyzed photo ATRP

We recently demonstrated that the remarkable swelling capability of CM resin facilitated the solid-phase synthesis of multi-block copolymers without the need for cumbersome and was-

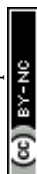




Fig. 4 Synthesis of protein-polymer hybrids (PPH) by heterogeneous CM-EY catalyzed photo-ATRP (A) BSA-PCBMA₂₂ and (B) CT-PCBMA₁₂. Reaction conditions: ([CBMA]/[Protein-I]/[CM-EY]/[CuBr₂]/[TPMA]/[TEOA]) = 200/1/0.005/0.2/0.6/0.6 in 1× PBS under green light irradiation (525 nm, 25 mW cm⁻²), CM-EY = 1.6 mg mL⁻¹.

teful purification steps between each block synthesis.⁸⁷ Here, by employing heterogeneous CM resin functionalized with ATRP initiators (CM-BIB), we facilitated the solid-supported synthesis of polymers *via* photo ATRP using CM-AM-EY as the photocatalyst. This system uniquely integrates both ATRP initiator moieties and the PC (*i.e.*, EY) onto a solid-support CM (Fig. 5). CM-BIB is composed of an ester linkage between hydroxy-functionalized CM and the ATRP initiator, while CM-AM-EY features an amide linkage between amine-functionalized CM and EY. The controlled ATRP process is governed by the interactions of the growing polymer chain end with a small-molecule catalyst (Cu^{II}Br₂/L), which can diffuse into the crosslinked network of the CM resin, thus ensuring high chain-end fidelity.⁹⁸

MA monomer was grafted from the CM-BIB in DMSO at room temperature using heterogeneous CM-EY as the photocatalyst and Cu/Me₆Tren as the co-catalyst, ethyl α -bromoisobutyrate (EBIB) was used as the homogenous sacrificial initiator in the system. After 60 minutes of irradiation of the polymerization reaction in capped SPE cartridge under green light LEDs (λ_{max} = 525 nm, 25.0 mW cm⁻²), the polymerization mixture was separated from the heterogeneous solid supports. ¹H NMR of the solution phase showed 50% conversion of [MA] monomer achieved. 20% TFA/H₂O was used to cleave the ester bond on CM-BIB releasing the grafted PMA. SEC analysis of PMA revealed that the polymerization pro-

ceeded with a high-control over molecular weight and dispersity in both solution and solid-phase (Fig. 5A).

To demonstrate one-pot synthesis of block copolymer on the solid support, the CM- grafted with PMA was thoroughly washed with DMSO and, without further purification, mixed with the next monomer [EA], Cu/Me₆Tren as the co-catalyst, and EBIB initiator. The polymerization mixture was irradiated for 60 min to graft PMA-*b*-PEA block copolymer from the CM resin. The acid-cleaved PMA-*b*-PEA block copolymer revealed clear shifting in the SEC trace towards high molecular weight ($M_{n,abs}$ = 15000, $\bar{D}$ = 1.26) without any shoulder, indicating successful chain extension on solid support was achieved (Fig. 5B).

Furthermore, utilizing an orthogonal linkage strategy to synthesize CM-BIB and CM-AM-EY enabled the recycling of this heterogeneous solid-support combination for subsequent ATRP reactions. The mixture of CM-grafted PMA and CM-AM-EY was subjected to acid hydrolysis (20% TFA/H₂O), thoroughly washed with DMSO, and re-functionalized with the ATRP initiator. During acid hydrolysis, CM-BIB releases the PMA, while CM-AM-EY remains unchanged. The re-functionalized CM-BIB resin and CM-EY mixture was successfully used for subsequent solid-phase photo ATRP, demonstrating the versatility of the solid-supported ATRP initiator and photocatalyst (Fig. 5C). However, SEC traces revealed a slightly lower molecular weight of cleaved PMA compared to solution-phase





Fig. 5 (A) Solid-phase synthesis of polymers catalyzed by heterogeneous CM-AM-EY (A) homopolymer of PMA grafted from CM-BIB, (B) block copolymer of PMA-*b*-PEA grafted from CM-BIB and (C) recycling of CM-BIB and CM-AM-EY for solid-phase polymer synthesis.

PMA, likely due to incomplete PMA cleavage from the previous step, resulting in sample contamination during recycling. This underscores the necessity of efficient TFA cleavage before recycling. Overall, this work highlights the ability to selectively cleave and reattach polymer chains from CM-BIB containing ester linkages while maintaining the amide linkages in CM-AM-EY intact.

Ir(III)-functionalized ChemMatrix (Ir@CM)

Encouraged by the successful photo ATRP achieved by CM-AM-EY, we sought to expand the scope of CM-based heterogeneous PC to other photocatalytic moieties such as transition metals. Therefore, we attempted to immobilize Ir(III) complexes through coordinative bonding and investigated the catalytic activity under blue light.

Photoluminescent cyclometalated iridium(III) complexes are promising photoactive compounds for their highly tunable excited state properties *via* ligand manipulation,^{99–102} as well as for their wide applications in sensors,¹⁰³ probing,¹⁰⁴ optoelectronics^{101,105} and photoredox catalysis.^{106–108}

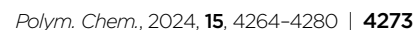
Particularly, heteroleptic $[\text{Ir}(\text{C}^{\wedge}\text{N})_2(\text{N}^{\wedge}\text{N})]^+$ complexes (where $\text{C}^{\wedge}\text{N}$ refers to 2-phenylpyridyl-type ligands, and $\text{N}^{\wedge}\text{N}$ refers to 1,2-diimine-type ligands) have been widely studied due to their enhanced photostability and long-lived triplet excited state caused by spin-orbital couplings and rapid intersystem crossing (ISC).¹⁰⁹ Thus, photoactive Ir(III) complexes have been used to catalyze various types of photoredox reactions, such as hydrogen evolution,^{107,110} water reduction,^{106,111} dehalogenation¹¹² as well as photoinduced polymerizations.^{113–115}

Streamlined “continuous-flow” synthesis of Ir@CM

The photoluminescent Ir@CM was obtained by functionalizing a commercially available hydroxy-terminated CM resin using a streamlined, “continuous-flow”-like setup in a solid-phase extraction (SPE) syringe, that enabled the multistep functionalization without tedious purification. The iodination of the –OH group was first performed based on a reported procedure,¹¹⁶ followed by the azidation¹¹⁷ and Cu-catalyzed azide-alkyne cycloaddition forming 2-triazol-pyridine-type ($\text{N}^{\wedge}\text{N}$) ligand.¹¹⁸ The removal of small-molecule side products and



To further validate the photophysical properties of these relatively unusual Ir^{III} complexes, the excited state properties



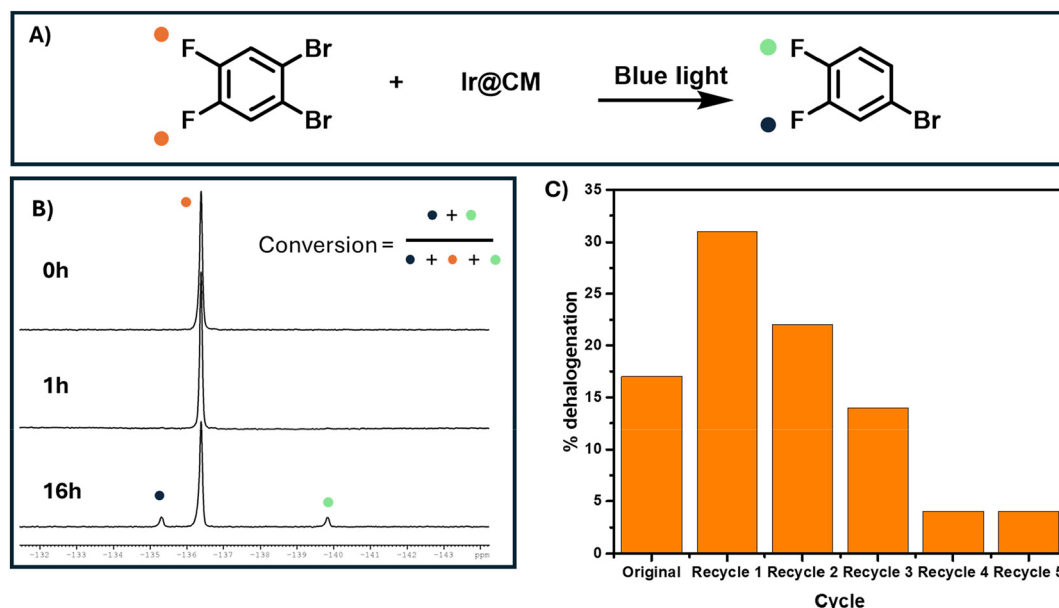


Fig. 7 (A) Reaction scheme of dehalogenation; (B) ^{19}F NMR of reaction mixture at different time points. Debromination is indicated by the appearance of two peaks at -135.25 ppm and -139.75 ppm. (C) Dehalogenation efficiency of pbtz-Ir@CM over numerous cycles determined by integration of peaks at -135.25 ppm and -139.75 ppm against 1,2-dibromo-4,5-difluorobenzene peak at -136.4 ppm.

of a model Ir complex containing the same C[^]N and N[^]N ligands were studied by density functional theory (DFT) calculations, which revealed a locally excited triplet state and explained the origin of the vibronic patterns observed in the emission spectra (Fig. S11[†]).

Investigation of catalytic performance

To probe the catalytic activity of pbzt-Ir@CM, we performed dehalogenation of 1,2-dibromo-4,5-difluorobenzene, which can be easily monitored *via* ^{19}F NMR. Following a previously reported method, pbzt-Ir@CM was used as a heterogeneous catalyst for debromination under blue light. After 1 hour, no change could be observed in the reaction solution (Fig. 7A). At 16 h, the peaks around -135.25 and -139.75 ppm emerged, indicating successful debromination of a single bromine (Fig. 7B). The second bromine proved harder to remove, as the loss of an electron-withdrawing group resulted in an increased oxidation potential for the remaining bromine. These peaks were then integrated against the original peak at -136.25 ppm to determine the dehalogenation efficiency of 17% (Fig. 7C).

The catalyst was then recycled five times until dehalogenation was no longer observed (Fig. 7C). After the first cycle, dehalogenation efficiency increased. This is explained by the pulverization of the pbtz-Ir@CM, increasing the surface area of the catalyst, allowing for more efficient dehalogenation. On subsequent cycles, a progressive decrease in dehalogenation efficiency was observed. This was mainly attributed to the leaching of the iridium complex, as the solutions showed slight discoloration after the reaction and photobleaching of the compound.

Comparison between covalent vs. non-covalent attachment of photocatalyst

In conclusion, this study demonstrates that the covalent linkage between the photocatalyst EY and the solid support (CM-EY) offers numerous advantages over non-covalent interactions (Ir@CM), which are more prone to disruption, potentially leading to catalyst leaching from the support. This enhanced stability significantly reduces the risk of catalyst detachment or deactivation during reactions, resulting in improved performance and longevity, as evidenced by the efficient recycling of CM-AM-EY for photo ATRP. Moreover, covalent attachment effectively prevents photocatalyst aggregation, thus minimizing photobleaching.

Furthermore, the precise positioning and orientation of the catalyst on the solid support enabled by covalent attachment may enhance its interaction with reactants, thereby improving catalytic efficiency and selectivity. In contrast, non-covalent interactions may lack the precise control necessary for optimal catalyst orientation and positioning. Covalent linkage also yields a more robust catalyst-support system, facilitating easier separation and recycling of the catalyst post-reaction. This aspect is crucial for cost reduction and waste minimization.

Moreover, covalent linkage offers greater flexibility for tailoring the properties of the catalyst-support system, such as adjusting loading density or introducing specific functional groups. This capability allows for fine-tuning of catalytic performance to meet the requirements of various reactions or applications. Overall, covalent linkage between a catalyst and a solid support provides a more stable, efficient, and versatile platform for heterogeneous catalysis compared to non-covalent bonding interactions.

Author contributions

The manuscript was written through the contributions of all authors. All authors have given approval to the final version of the manuscript.

Data availability

The authors confirm that the data supporting the findings of this study are available within the article and/or its ESI.†

Conflicts of interest

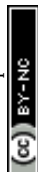
The authors declare no competing interests.

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