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Cite this: *Mater. Chem. Front.*,
2019, 3, 762

Pyrene-based aggregation-induced emission luminogens and their applications†

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"Unity is force" – Aesop. It is a common phenomenon that traditional chromophores exhibit high fluorescence in dilute solutions, yet luminescence is quenched at high concentrations or in the aggregate state, *i.e.* "aggregation-caused quenching" (ACQ). Tang reported the unusual photophysical observation that luminogens can exhibit weak or no fluorescence in solution, yet they are highly emissive in the aggregate or solid state; this is defined as aggregation-induced emission (AIE). The discovery of AIE helped solve the ACQ effect in traditional luminophores. Pyrene is an important polycyclic aromatic hydrocarbon (PAH), which exhibits very different photophysical behavior in solution *versus* the aggregate state, and the ACQ effect has played a dominant role in pyrene chemistry. The ACQ effect is harmful for some practical applications and is a challenge in organic light-emitting diodes (OLEDs) and light-emitting electrochemical cells, for which the effect is more severe in the solid state. Thus, how to overcome the ACQ effect observed in pyrene chemistry still remains a challenge. In this review, we discuss how following basic AIE mechanisms such as the restriction of intramolecular motion (RIM), excited-state intramolecular proton transfer (ESIPT), and twisted intramolecular charge transfer (TICT), can transform pyrene-based ACQ luminogens to AIE luminogens with excellent optical properties. Furthermore, prospective applications of pyrene-based AIEgens are discussed, as is the potential for designing new organic functional materials.

Received 11th February 2019,
Accepted 18th March 2019

DOI: 10.1039/c9qm00090a

rsc.li/frontiers-materials

1. Introduction

1.1. Chemistry of pyrene and its derivatives

Pyrene is a representative member of the polycyclic aromatic hydrocarbons (PAHs), and consists of four fused benzene rings with extensive π -electron delocalization. Since Weitzenböck first synthesized pyrene **1** in the laboratory in 1913, pyrene **1** (Fig. 1) and its derivatives have attracted much attention from scientists, primarily due to their blue emissive property, good hole-transporting ability, and their application in organic electronics, such as organic light-emitting diodes (OLEDs), light-emitting electrochemical materials, solar cells *etc.*^{1–3} Salient features of pyrene are that its π -system does not require protection during functionalization,⁴ its moderately long fluorescence lifetime, and excimer formation.⁵ However, there are two major drawbacks of

pyrenes when used in organic fluorescent materials, (i) the absorption and emission of the pyrene fluorophore lies in the near-ultraviolet (UV) region, whereas a good optoelectronic device emits in the visible region and (ii) excimer formation occurs in concentrated solutions or in the solid state, which decreases or quenches the fluorescence intensity, with a large red-shifted emission.^{6–8} This common phenomenon is defined as aggregation-induced quenching (ACQ). For example, PAH-based luminogens like pyrene and its derivatives emit deep blue emission with high quantum yields at low concentrations, which has limited its application as an efficient solid emitting material.^{2,9}

The intrinsic properties of pyrene alone do not make it a good candidate for light-emitting devices; however, the ACQ effect of pyrene has been overcome by introducing bulky groups with long alkyl chains which enhances the fluorescence quantum efficiency in OLEDs and also reduces the aggregation of the final material.² For example, a 1,3,6,8-substituted pyrene-based conjugated molecule such as 1,3,6,8-tetrakis(4-butoxyphenyl)pyrene exhibits an active emitting layer in simple solution-processed OLEDs producing deep blue emission (CIE = 0.15, 0.18) and maximum efficiencies and brightness levels of 2.56 cd A^{−1} and 5000 cd m^{−2}.¹¹

Experimentally, the planar structure of pyrene prefers to adopt a close-packing form in the aggregation state. The π – π stacking interaction prompts the formation of detrimental species such as

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† This review is dedicated to Emeritus Professor Takehiko Yamato (Saga University, Japan) on the occasion of his 66th birthday.

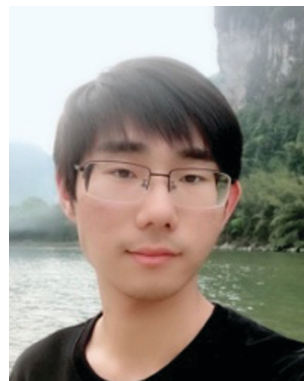
excimers, resulting in the observed ACQ effect. Pyrene-based highly efficient emitting materials can be obtained by introducing various substituents at the 1-, 3-, 6-, and 8-positions of the pyrene core.² Generally, these four positions are significantly more active than the other positions (the 4-, 5-, 9-, and 10-positions or the 2- and 7-positions), and it is not easy to directly substitute at these latter positions *versus* the 1-, 3-, 6-, and 8-positions.⁹

Theoretically, the energy of the 1-, 3-, 6-, and 8-positions are lower by 8.8 kcal mol⁻¹ than the 2- and 7-positions, and the substitution order follows the trend 1- > 8- > 6- > 3-.¹² Pyrene can be easily modified at the 1-, 3-, 6- and 8-positions to form mono-, di-, tri-, and tetra-substituted derivatives¹³ in a manner that can hinder π - π interactions to effectively suppress the excimer formation.¹⁴ The introduction of bulky groups at the 1-, 3-, 6- and



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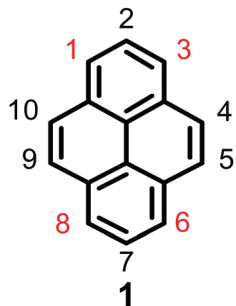


Fig. 1 The molecular structure of pyrene 1.

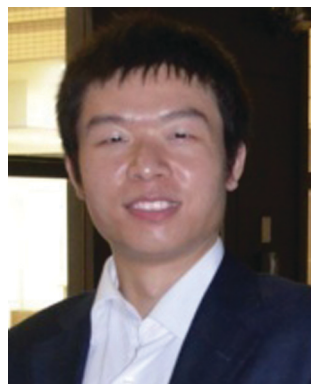
8-positions of pyrene leads to a shift in the main absorption and emission bands into the visible region due to the extension of the π -conjugation system. On the other hand, as the nodal plane of pyrene passes through its carbon atoms at the 2,7-positions in the highest occupied molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO), substituents at these positions will interact weakly with the central core.¹⁵ For example, the uridine ring exhibited a weak electronic interaction with the pyrene core when it was substituted at the 2-position instead of the 1-position.¹⁶ Thus, our group has previously focused on synthetic methodologies for the functionalization of the pyrene core to suppress the ACQ effect based on its special electronic structure, and numerous novel synthetic routes for the functionalization of the pyrene core have been explored. Examples include *tert*-butylation, electrophilic substitution, formylation, halogenation, oxidation, and borylation. Moreover, other research groups have also attempted to explore a wide variety of methods including

physical and engineering approaches and processes, but the results to date are far from satisfactory.¹⁷

1.2. Pyrene-based aggregation-induced quenching luminogens

The design and synthesis of efficient, stable pyrene-based fluorescent materials has been an active topic in the area of OLEDs.^{10b,18} However, the π -aggregates/excimers in concentrated solution and the solid state,¹⁹ lead to the formation of π -aggregates/excimers and subsequent quenching of fluorescence with low photoluminescence (PL) quantum yields (Fig. 2).^{20,21} As a result, much effort has been devoted to improving the photophysical properties of pyrene.^{1a,22} Thus far, the pyrene derivatives that have been reported still present some degree of fluorescence quenching in the solid state.^{6,11} Therefore, the preparation of compounds based on pyrene that emit efficiently in the solid state remains an attractive objective.

Research to modulate the photophysical properties of pyrenes to enhance fluorescence both in the solid and (or) in thin film is currently being conducted.²³ Using synthetic methods for the functionalization of the pyrene core, 1-substituted-, 1,3-disubstituted-, 2,7-di-substituted-, 1,3,6,8-tetrakis-substituted-, 1,3,5,9-substituted-, and 4,5,9,10-substituted-pyrenes have been explored. However, all of above-mentioned compounds are ACQ compounds, because of the low quantum yields in the solid (film) state compared with those in solution. For example, 1,3,6,8-tetraphenylpyrene (TPPY; Fig. 2) is highly fluorescent (Φ ca. 0.9) in solution, and the organic light-emitting field-effect transistor devices (OLEFET) based on TPPY have been shown to exhibit electroluminescence (EL) with an external quantum efficiency of only 0.5% due to aggregation.^{6,24} Thus, conventional dyes suffering from the ACQ effect may not be suitable for constructing highly efficient non-doped OLEDs. However, materials with non-planar propeller-shaped molecular structures exhibiting aggregation-induced emission (AIE) characteristics may be an ideal choice for non-doped OLEDs.



Xing Feng

Xing Feng received his PhD (2013) from Saga University (Japan) under the supervision of Prof. Takehiko Yamato. Then, he joined the School of Beijing Institute of Graphic Communication (BIGC) as a lecturer. He was an established academic researcher in BIGC and held a position at the Ministry of Science and Technology of China, being in charge of the administration of the National Program for International Science and Technology Cooperation in

2015. In 2016 he joined Professor Ben Zhong Tang's group in the Department of Chemistry at the Hong Kong University of Science and Technology (HKUST) as postdoctoral fellow. In 2017, he moved to the School of Materials and Energy, Guangdong University of Technology (GDUT) in Guangzhou as an associate professor. His research interests include the development of novel fluorescence materials and fluorescent bioprobes for application in organic electronics. He was awarded with the Chinese Government Award for Outstanding Self-Financed Students Abroad by the China Scholarship Council in 2013.

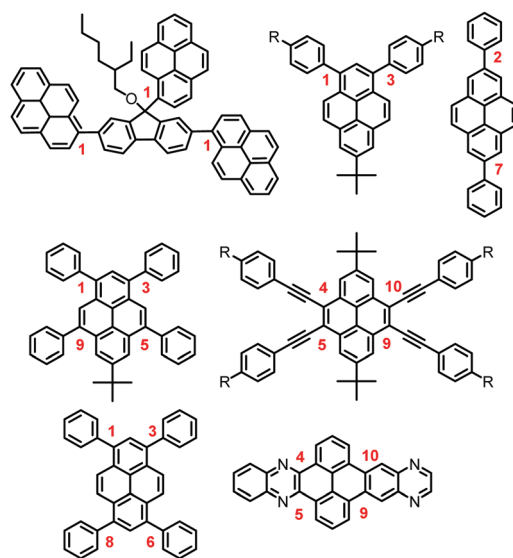


Fig. 2 Chart of the molecular structures of pyrene-based blue emitters with the ACQ feature.^{19,20}

1.3. Aggregation-induced emission

It is essential for organic luminescent materials to exhibit excellent optoelectronic properties (such as high fluorescence, good electronic/hole mobility *etc.*) for potential application in organic electronics, especially in the solid state or aggregated state. Tang *et al.* observed an abnormal photophysical phenomenon in hexaphenylsilole (HPS) in 2001 (Fig. 3), whereby weak emission was present in solution but enhanced emission in the solid state.²⁵ This phenomenon was termed “aggregation-induced emission” (AIE), and the discovery of AIE solved the ACQ effect of traditional luminophores.²⁵ Recently, the AIE concept has been widely accepted in the scientific community, and more than 40 000 papers on AIE research have now been published, including on siloles,²⁶ tetraphenylethylene (TPE),²⁷ 9,10-bis(*p*-dimethylaminostyryl)-anthracene (9,10-MADSA),²⁸ 1-cyano-*trans*-1,2-bis-(40-methyl biphenyl)ethylene (CN-MBE),²⁹ 2,5-diphenyl-1,4-distyryl-benzene (DPDSB) derivatives,³⁰ diphenyldibenzofulvene (DPDBF) derivatives,³¹ conjugated polymers,³² borondipyrromethene (BODIPY) derivatives,³³ triazoles,³⁴ and others.^{35,36}

Different hypotheses have been put forward for AIE mechanisms, such as J-aggregates, excited-state intramolecular proton transfer (ESIPT), and twisted intramolecular charge transfer (TICT) *etc.*³⁷ The restriction of intramolecular motion (RIM) was widely accepted to explain the AIE mechanism. Take hexaphenylsilole (HPS) as an example – the six peripheral phenyl rotors were dynamically rotated relative to the central silole, around the C–C or C–Si bond. The rotation of the aromatic rotors generally destroys the excitons in a non-radiative mode in the solution state. Upon aggregate formation, the emission intensity of the AIEgens would be enhanced by the restricted intramolecular rotation (RIR), which blocks the non-radiative pathway for the excitons to decay with the activated radiative transition, thus the radiative decay of the excited state becomes dominant, rendering the luminogens emissive. As a consequence, the fluorescence quantum yields of siloles in aggregates is two orders of magnitude greater than in solution due to the restriction of RIR of the chromophoric compound.²⁵

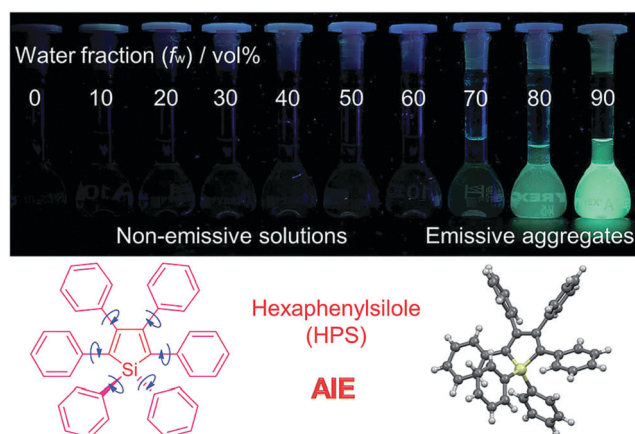


Fig. 3 Fluorescence photographs of solutions and suspensions of hexaphenylsilole (HPS; 20 mM) in tetrahydrofuran (THF)/water mixtures with different fractions of water. Reproduced from ref. 25d. Copyright (2014) Wiley-VCH.

1.4. Jablonski diagram

A molecule in the excited state can decay in several ways, as shown in the Jablonski diagram (Fig. 4).³⁸ Relaxation to the ground state from these excited states occurs by a number of mechanisms that are either radiation-less, in that no photons are emitted, or involve the emission of a photon. The most likely pathway by which a molecule relaxes back to its ground state is that which gives the shortest lifetime for the excited state, of the order of 10^{-15} s.³⁸

By a combination of internal conversions and vibrational relaxations, a molecule in an excited electronic state may return to the ground electronic state without emitting a photon. A final form of radiation-less relaxation is an intersystem crossing in which a molecule in the ground vibrational energy level of an excited electronic state passes into a high vibrational energy level of a lower energy electronic energy state with a different spin state. For example, an intersystem crossing between a singlet excited state (S_1), and a triplet excited state (T_1). Fluorescence ($S_1 \rightarrow S_0 + h\nu$) occurs when a molecule in the lowest vibrational energy level of an excited electronic state returns to a lower energy electronic state by emitting a photon.

Phosphorescence ($T_1 \rightarrow S_0 + h\nu$) is most favorable for molecules that have $n \rightarrow \pi^*$ transitions, which have a higher probability for an intersystem crossing than do $\pi \rightarrow \pi^*$ transitions. Delayed fluorescence results from two intersystem crossings, first from the singlet to the triplet, and then back to the singlet state ($S_1 \rightarrow T_1 \rightarrow S_1 \rightarrow S_0 + h\nu$).

The intensity of fluorescence or phosphorescence I , is proportional to the amount of radiation from the excitation source that is absorbed and the quantum yield for fluorescence.

$$I = k\Phi(A_0 - A) \quad (i)$$

where k is a constant accounting for the efficiency of collecting and detecting the fluorescent emission. From Beer's law we know that

$$A/A_0 = 10^{-\epsilon bc} \quad (ii)$$

where c is the concentration of the fluorescing species. Solving eqn (ii) for A and substituting into eqn (i) gives, after simplifying

$$I = k\Phi A_0(1 - 10^{-\epsilon bc}) \quad (iii)$$

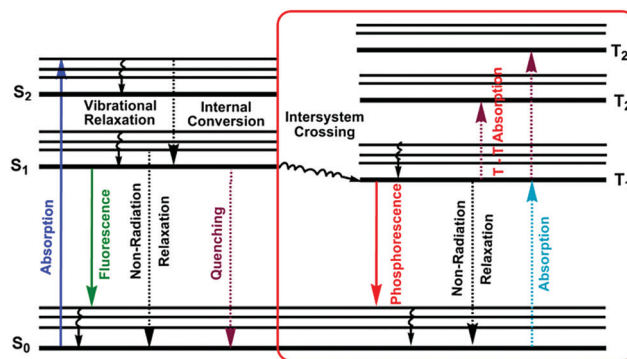


Fig. 4 A Jablonski diagram of the relevant photophysical processes in luminescence materials.

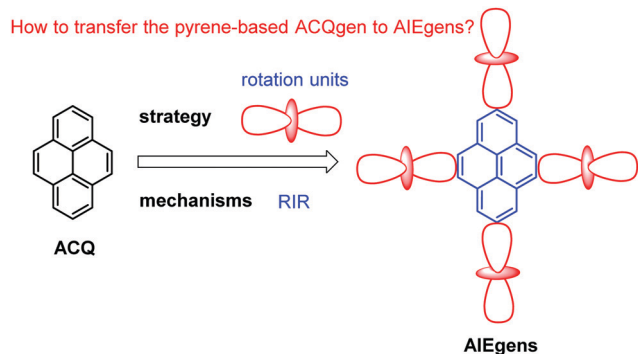


Fig. 5 A schematic diagram of pyrene-based AIEgens possessing bright emission.

For low concentrations, where εbc is less than 0.01, this equation simplifies to

$$I = 2.303k\Phi A_0 \varepsilon bc \quad (\text{iv})$$

According to the Jablonski diagram, upon aggregate formation, the emission of AIEgen is induced or revitalized by the specific effects of the RIM and the highly twisted molecular conformation that destroys the intermolecular π - π stacking interactions.^{25,26} We will first carefully elaborate on the working principle of AIE and then demonstrate how to transform traditional ACQgens into new AIEgens, followed by the presentation of a series of new AIEgens derived from the RIM principle (Fig. 5). Finally, we will recap the research efforts devoted toward the exploration of technological applications of the new pyrene-based AIEgens. To understand the mechanisms of AIE phenomena is of great importance in the quest for fundamental knowledge of photophysics, and just as crucial, it will guide our endeavors to design novel pyrene-based AIE luminogens, and to explore practical applications and technological innovations. Since the arrival of the AIE concept in 2001, researchers have shown a persistent thirst for learning what are the real mechanistic causes of the AIE phenomena.^{25,26} Herein, we will summarize new pyrene luminogens, and more importantly, discuss a feasible method to realize the transformation of pyrene-based ACQ luminogens to AIE. The excellent optical properties of pyrene-based AIEgens for potential applications in organic semiconductor devices in biochemistry will be discussed, as well as recent developments in the area of chemisensors.

2. Pyrene-based aggregation-induced emission luminogens

2.1. Pyrene core AIEgens

Since the emission of pyrene **1** is either weak or quenched in the solid state due to its planar structure with a large π -conjugation system, an effective approach to overcome this is the incorporation of bulky substituents or a spiro-type framework at the periphery of the pyrene core in order to inhibit the π - π interactions, leading to a highly emissive solid state. Arylsilyl substituted pyrene derivatives **2** and **3** were synthesized *via* Suzuki–Miyaura/Sonogashira–Hagihara cross-coupling reactions to afford the

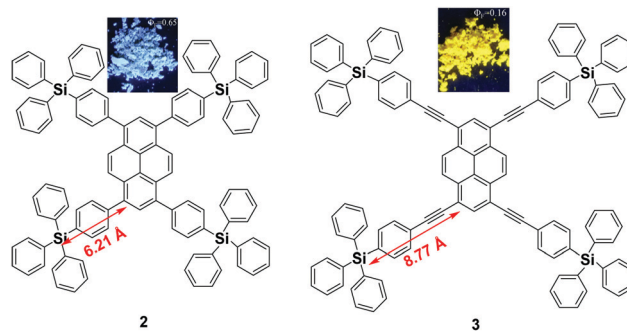


Fig. 6 Structures of **2** and **3** calculated via DFT methods. Reproduced from ref. 39.

corresponding silyl-terminated group at the 1-, 3-, 6- and 8-positions. Triphenylsilylphenyl substituted pyrene **2** exhibited intense emission at 440 nm in the film with a high fluorescence quantum yield ($\Phi_f = 0.65$); whereas the triphenylsilylphenyl-ethynyl substituted pyrene **3** showed an excimer emission with low Φ_f both in THF solution (0.50) and as a film (0.16).³⁹ Despite the introduction of the bulky tetraphenylsilyl ethynyl groups in **3**, the pyrene moieties can clearly still interact in an orthogonal orientation. Density functional theory (DFT) geometry optimization predicts that the incorporation of the coplanar phenylethynyl substituents in the structure of **3** results in the Si atoms being 8.77 Å from the 1-, 3-, 6-, and 8-carbon atom positions, rather than 6.21 Å, as is the case in the structure predicted for **2** (Fig. 6).³⁹

The longer distance between the Si atoms and the 1-, 3-, 6-, and 8-positions on the outer perimeter of the pyrene π -system, probably enables significant π - π interaction to occur between neighboring pyrenes. Although the pyrene possesses a large π -conjugation planar structure, according to the RIR mechanism, it is possible to convert the pyrene into a highly efficient pyrene-based AIEgen by introducing a twisted fragment, such as TPE or triphenylethylene (TriPE) units *etc.* Zhao first introduced the TPE units at the pyrene core as peripheral ornaments affording TTPEPy **4** possessing a novel aggregation-induced enhanced emission (AIEE) effect (Fig. 7). It emitted sky-blue emission (483 nm) with efficient PL ($\Phi = 70\%$) in the solid state, and excellent thermal stability ($T_d = 485^\circ\text{C}$).⁴⁰ The temperature-dependent PL spectrum was measured over the range 300 K to 77 K, and the fluorescence lifetime gradually lengthened from

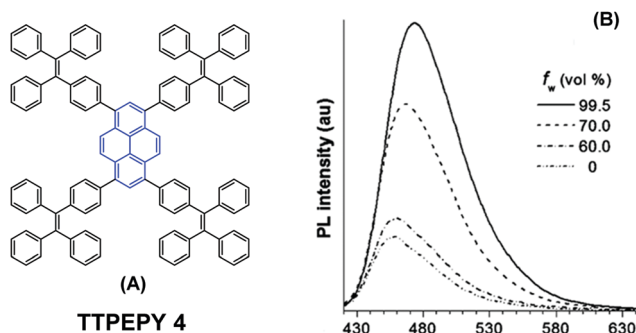


Fig. 7 The structure of TTPEPy **4** and emission spectrum of TTPEPy **4** in THF with increasing percentages of water. Reproduced from ref. 40.

0.25 ns at 300 K to finally 1.29 ns at 77 K. This was due to the fact that the intramolecular rotations of the TPE blades were greatly restricted at the cryogenic temperature and the accompanying RIR process helped enhance the PL of TTPEPy 4 in the frozen solution.

In addition, the construction of efficient non-OLEDs using TTPEPy 4 as a blue emitter, which display an excellent current efficiency (η_c), power efficiency (η_p), and maximum external quantum efficiency (η_{ext}) up to 12.3 cd A⁻¹, 7.5 lm W⁻¹ and 4.95%, respectively has been reported.⁴⁰ Furthermore, to investigate the relationship between the molecular conformation, the emission color and the AIE properties, Li *et al.* designed and synthesized two pyrene-based luminogens, Py-4MethylTPE (5) and Py-4mTPE (6) (Fig. 8). The differing behavior is due to the linkage mode at the *meta*-position which can twist the molecular structure and provide a new stereoelectronic effect (Fig. 8). Both compounds exhibit AIE and AEE effects, although they have minor differences in their structures;⁴¹ for example the torsion angle of 5 between the pyrene core and the TPE periphery is larger than that of 6 due to the introduction of the methyl groups. On the other hand, use of *meta* linkages on the TPE moieties shortens the π -conjugation length in comparison with the *para* linkages and results in the increased rotational barriers. The *meta* linkage is more effective at controlling the twisted conformation and generating enhanced molecular rotation. Given this, the compound 6 emits a more enhanced blue emission at 445 nm compared with 5, which shows an emission peak at 459 nm in the aggregation state. Importantly, when using compound 6 as emitter, the OLED device exhibits favorable EL properties with a lower turn-on voltage (3.5 V) and better OLED performance, with values of $\eta_{C,\text{max}}$, $\eta_{P,\text{max}}$, and $\eta_{\text{ext,max}}$ at 4.02 cd A⁻¹, 3.08 lm W⁻¹, and 2.5%, respectively.^{41a}

It seems that it is an efficient approach to achieve pyrene-based AIEgens through integrating pyrene with TPE units. Tao's group has presented a pyrene-based fluorescent material, namely 1-[4-(2,2-diphenylvinyl)phenyl]pyrene (PVPP 7), which was synthesized *via* the Suzuki coupling reaction with TriPE units.^{20b} Interestingly, compound 7 shows a dramatic enhancement of fluorescence intensity with clear AIEE characteristics. The single crystal X-ray diffraction analysis indicated that the two terminal phenyl rings in PVPP 7 are almost perpendicular, with a twist angle of 89.3°, and that several intramolecular C-H... π interactions are present which can play a significant role in suppressing the

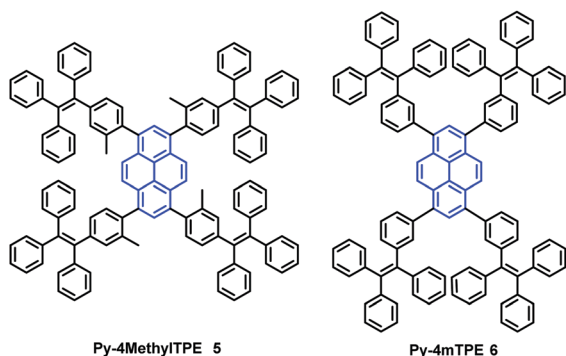


Fig. 8 Structures of Py-4MethylTPE 5 and Py-4mTPE 6.

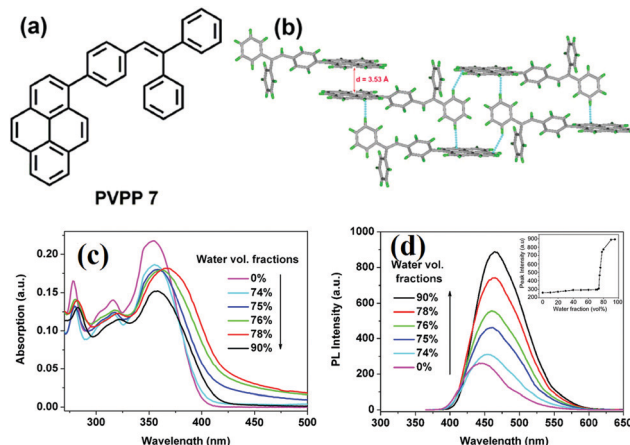


Fig. 9 (a) The molecular structure of PVPP 7 and (b) the crystal packing structure PVPP 7. (c) UV-visible absorption spectra of PVPP 7 in THF/water mixtures with different volume fractions of water. (d) PL spectra of PVPP 7 in THF/water mixtures with different volume fractions of water. The inset shows the relationship between the PL peak intensity of PVPP 7 and the water fraction in THF. Reproduced from ref. 20b.

fluorescence quenching of the pyrene units in the solid state (Fig. 9d). The free twisting motion of the 1,1-diphenylvinyl motif is restricted by the intra- and intermolecular C-H... π hydrogen bonds. Thus, the non-radioactive decay channel is closed. More importantly, the molecular structure of PVPP 7 is significantly distorted, and the twisted geometry can prevent the strong π - π interactions producing intense fluorescence in the solid state. Li *et al.* first introduced the twist terminal units at the 2,7-positions of pyrene to access the pyrene-based AIEgens 8–13 (Fig. 10). In this case, the AIE building blocks become more and more twisted, and large blue-shifted emission is achieved on moving from Py-2mTPE 9 (460 nm) to Py-2F 12 (456 nm) and to Py-2TP 13 (444 nm).^{41b}

Py-2 NTF 11 exhibited a blue-shifted absorption compared with Py-2F 12, owing to the weak conjugation at the 9,9'-positions in the fluorene as well as the severely twisted conformation of this linkage mode. There was little difference between Py-2NTF 11 and Py-2TF 10 in the UV absorption spectra due to the non-conjugated 9,9'-positions of fluorene. Py-2pTPE 8 and Py-2mTPE 9, which consisted of typical AIE units of tetraphenylethene, show AIE characteristics (non-emissive in THF solution but a strong fluorescence in the aggregation state).

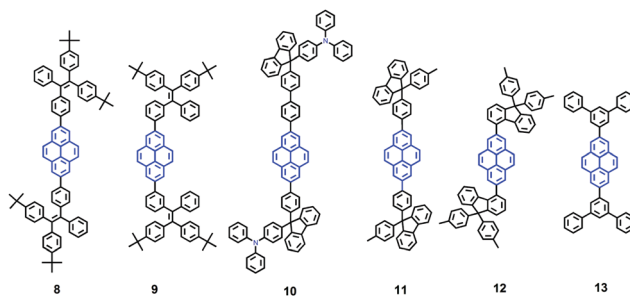


Fig. 10 Structures of Py-2pTPE 8, Py-2mTPE 9, Py-2TF 10, Py-2NTF 11, Py-2F 12, and Py-2TP 13.

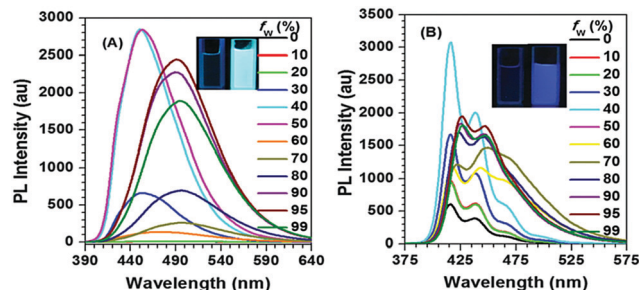


Fig. 11 PL spectra in THF/H₂O mixtures with different water fractions: (A) Py-2pTPE **8**, concentration (mM): 12.0; excitation wavelength 330 nm; (B) Py-2TP **13**, concentration (mM): 14.0; excitation wavelength 300 nm. Inset: Photos of Py-2pTPE **8** and Py-2TP **13** in THF/H₂O mixtures (f_w = 0 and 99%) taken under illumination by a 365 nm UV lamp. Reproduced from ref. 41b.

Interestingly, for Py-2pTPE **8** (Fig. 11), as the water fraction was varied over the range 30–50%, it exhibited a deep blue emission which peaked at about 450 nm. The luminescence intensity increased gradually, and reached a maximum at 50% water fraction. However, when the water fraction increased to 60%, only very weak luminescence was observed, and its emission peak was red-shifted to about 494 nm. This abnormal phenomenon originated from the transition of the crystalline state to an amorphous state.^{41b}

More recently, the group of Tang reported a new strategy to achieve white-light emission from single tetraphenylethylene substituted pyrenes (TPE-Pys) **14–16** with AIE characteristics (Fig. 12).⁴² These molecules exhibited tuneable emission from blue (474 nm) to yellow (531 nm) both in solution and in the solid state. Specially, **14c** and **16** emitted a dual fluorescence at around 436 nm and 538 nm in a water fraction of 80%, which resulted in white-light with a CIE of $x = 0.30$, $y = 0.41$.⁴² These compounds emit white-colored fluorescence in THF/water mixtures. Indeed, these are rare examples of white-light emission achieved from a single AIE-active molecule *via* the control of its aggregated state emission by tuning the composition of the solvent mixture.

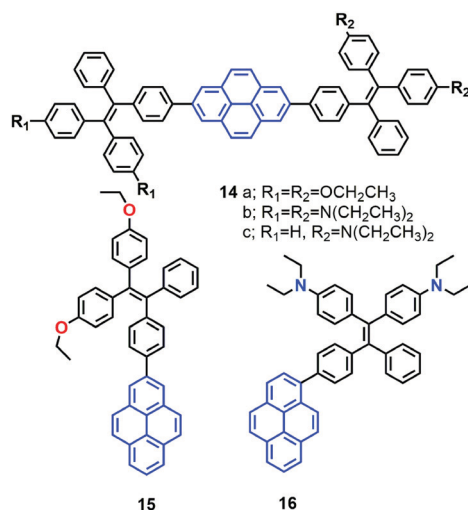


Fig. 12 Structures of TPE-Pys **14–16**.

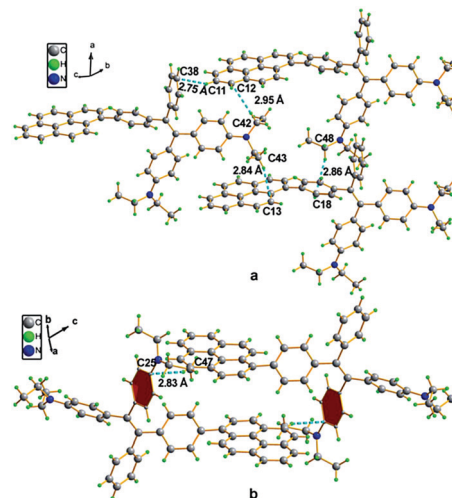


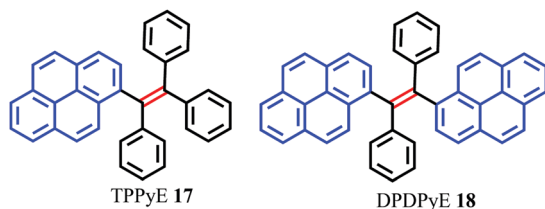
Fig. 13 Single crystal structure of **16**: (a) face-to-face pattern of pyrene moieties separated by TPE units and (b) head-to-tail packing mode with multiple C–H... π interactions with a distance of 2.83 Å. Reproduced from ref. 42.

A single crystal X-ray diffraction study indicated that compound **16** adopts a twisted conformation and revealed that the torsion angle between the terminal pyrene ring and the benzene ring of the TPE is 48.6° (Fig. 13). The molecular motion of the TPE moiety is suppressed by multiple intramolecular C–H... π interactions with distances of 2.75–2.95 Å. The pyrene rings are arranged in a head-to-tail fashion and are separated by the TPE units at a large distance of 8.08 Å. This prevents their π – π stacking which would quench the light emission. Multiple C–H... π interactions are formed in the crystal lattice which effectively restrict the phenyl rings of TPE from undergoing intramolecular motions. The non-radiative decay pathways are blocked to allow **16** to emit intense light in the aggregated state.⁴²

2.2. TPE-like pyrene-based AIEgens

Tetraphenylethylene (TPE) has attracted much attention because it possesses advantages such as efficient solid-state emission, facile synthesis, and ease of functionalization.^{43,44} More importantly, TPE is a promising candidate to convert conventional ACQ chromophores into effective AIEgens with excellent photoluminescence properties.⁴⁵ Given their outstanding AIE properties, much effort has been devoted to developing TPE-like based AIEgens which are expected to have large π -conjugation as well as enhanced emission behavior in the solid state. The main approach to obtaining efficient solid photoluminescence is to replace the phenyl rings of TPE by PAHs which introduce increased π -conjugation into the system. To date, several TPE-based AIEgens with notable solid-state emission have been developed.⁴⁶ Researchers have speculated whether a tetra(polycyclic aryl)ethene with excellent solid-state emission can be obtained by carefully selecting PAHs as substituent groups. These PAHs not only have an enlarged π -conjugated system, but also have effective steric hindrance to restrict the formation of π -stacking.

To verify this assumption, scientists have utilized many PAHs such as pyrene, fluorene, naphthalene, biphenyl *etc.* to assemble

Fig. 14 Structures of **17** and **18**.

tetra(polycyclic aryl)ethenes.⁴⁷ Zhao *et al.* reported two pyrene-substituted ethenes, namely TPPyE **17** and DPDPyE **18** (Fig. 14), where the phenyl ring of TPE was replaced by a pyrene ring. Unexpectedly, both the TPE-like pyrene-based dyes exhibited clear AIE features and were weak emitters in solution, but became strong emitters when aggregated in the condensed phase.^{46a} The crystal packing indicated that the intramolecular rotation was suppressed by π - π stacking and multiple C-H $\cdots\pi$ interactions. These structural features block the non-radiative energy decay channels, and allow for intense emission in the solid state.

Using TPPyE **17** and DPDPyE **18** as high-efficiency emitters, green OLEDs were fabricated with maximum luminance and power, current, and external quantum efficiencies of 49 830 cd m⁻², 9.2 lm W⁻¹, 10.2 cd A⁻¹ and 3.3%, respectively. To investigate the effect of substituent size on molecular rotation, Ni *et al.* reported tetrakis(4,5,9,10-tetrahydropyren-2-yl)ethene **19**, which exhibits clear AIE characteristics.⁴⁷ The fluorescence quantum yield of tetrakis(4,5,9,10-tetrahydropyren-2-yl)ethene **19** in a solid film was up to 74.1%. A crystal structure analysis indicated that no π -stacking interactions are present (Fig. 15). This study reveals that when constructing AIE molecules using substituent groups with expanded π -conjugation, the introduction of appropriate steric hindrance on the substituent groups can effectively restrict the formation of excimers. Meanwhile, hydrogen atoms from the ethylidene groups of compound **19** tend to interact with the π -cloud center of the closest two phenyl rings from the adjacent molecules, and these multiple C-H $\cdots\pi$ interactions rigidify the molecular structure and restrict the rotation of **19**, thereby making a valuable contribution in avoiding the formation of excimers. Thus, compound **19** exhibited high emission in the solid state with a high quantum yield (74.1%) compared with that observed in THF solution (1.25%).⁴⁷

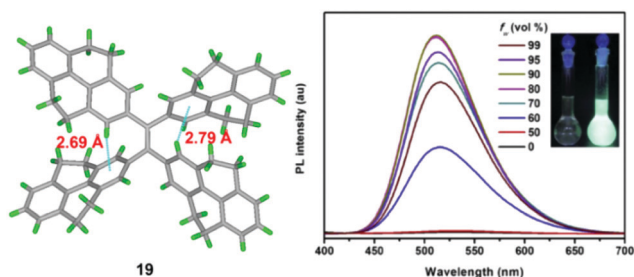
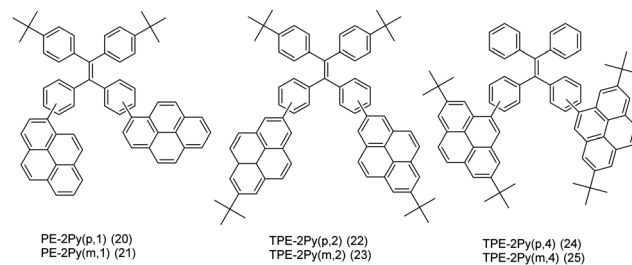


Fig. 15 Crystal structure of **19** and PL spectrum of compound **19** in THF/water mixtures with different water fractions (f_w). Inset: Photos of compound **19** in THF/water mixtures ($f_w = 0$ and 90%) under UV lamp illumination. Excitation wavelength: 370 nm. Reproduced from ref. 47.

Fig. 16 Structures of **20–25**.

On the other hand, it is an efficient approach to construct AIEgens by expanding the π -conjugation of the ethene derivatives using pyrene units. Li *et al.* further published a series of efficient TPE-like pyrene-based blue materials **20–25** with an AIE building block acting as the core and pyrene as a peripheral group. These compounds showed enhanced hole-transporting ability, achieving a good interaction between pyrene substituents (Fig. 16).⁴⁸

The effect of substituent position on the emission color and electroluminescence was investigated by modifying the linkage modes of the TPE (*para* to *meta*-position) and pyrene units (1- to 2- to 4-position). All of these pyrene derivatives exhibited apparent AIE characteristics and comparable EL performances in their OLED devices. When the pyrene units were attached at the *meta*-position of TPE, the PL emissions were tuned to deep blue (452 nm) in the solid state,⁴⁸ with that at the pyrene 4-position producing a better deep blue emitter than the 1- or 2-substituted pyrenes. Using TPE-2Py(*p*,4) **24** as both a hole-transporting layer and emitting layer, the resulting blue OLED device (CIE coordinate: $x = 0.16$, $y = 0.21$) exhibited excellent EL performance, with a current efficiency and external quantum efficiency up to 4.66 cd A⁻¹ and 2.79%, respectively.⁴⁸ Furthermore, Yanget *et al.* reported the highly efficient blue emitter TPE-4Py **26** with an AIE effect, which is achieved by a combination of a twisted tetraphenylethene (TPE) core and planar pyrenes on the periphery.⁴⁹ Its EL efficiency was enhanced in doped OLEDs, at high doping concentration (50%), with a current efficiency up to 4.9 cd A⁻¹ at 484 nm (Fig. 17). The twisted conformation of **26** led to weak

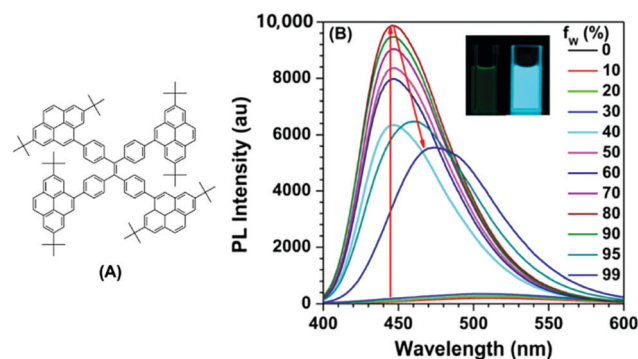


Fig. 17 (A) Molecular structure of TPE-4Py **26**; (B) PL spectra of TPE-4Py **26** in THF/H₂O mixtures with different water fractions. Concentration: 10 μ M; excitation wavelength: 330 nm; inset: photographs of TPE-4Py **26** in THF/H₂O mixtures ($f_w = 0\%$ and 90%) taken under illumination by a 365 nm UV lamp. Reproduced from ref. 49.

conjugation as well as blue emission. Furthermore, the π - π stacking was effectively restricted, which resulted in an abnormal AIE effect. The optimized structure revealed a heavily twisted conformation with a torsion angle of about 50° for the TPE, while the dihedral angle between the pyrene and the adjacent phenyl ring was as wide as 60° .⁴⁹

2.3. Non-typical pyrene-based AIEgens

A study of vinylene-based materials revealed that the vinylene bond geometry plays an important role on the emission.⁵⁰ For example, the *cis*-isomer in poly(*p*-phenylenevinylene) (PPV) molecules generally possesses weak luminescence.⁵¹ However, random placement of *cis*-olefin linkages along the PPV chain significantly increases the luminescence efficiency. For vinylene-based chromophores with an AIE or AIEE property, most tend to adopt the *trans*-isomer conformation.⁵¹ According to the RIR mechanism, the C=C bond can also play a significant role when constructing AIE molecules. For example, 1-[(9-anthracenyl)vinyl]pyrene (AVP)-based AIEgens *trans*-AVP 27 and *cis*-AVP 28 are presented in Fig. 18, and although the large planar structure of the pyrene and anthracene are both ACQ active, the anthracene ring was connected to a pyrene unit *via* a vinyl bridge and shows a clear AIEE characteristic with the two isomer conformations in the crystalline state.⁵²

The quantum yield of *trans*-AVP 27 and *cis*-AVP 28 are 54.0% and 36.0% in the solid state *versus* in solution 15.0% and 17.0%, respectively. The interacting adjacent molecules in crystals of *trans*-AVP 27 adopt both edge-to-face and face-to-face arrangements. In contrast, in crystals of *cis*-AVP 28, the adjacent molecules are inclined to adopt an edge-to-face arrangement and the molecular packing plays an important role in the solid emission efficiency.⁵² Similarly, Li *et al.* systematically investigated the substituent position and the vinylene bond geometry on the AIE features and the fluorescence solvatochromism. The structure of the (9-anthryl)vinyl(1-pyrenyl)-vinylbenzene isomers 29–32 are presented in Fig. 19.⁵³

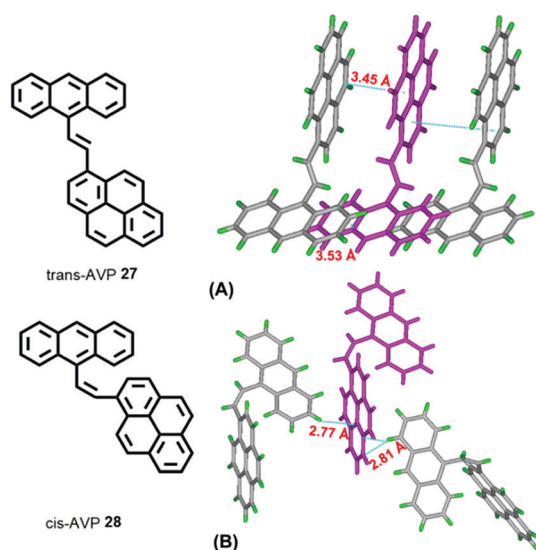


Fig. 18 Structures of (A) *trans*-AVP 27 and (B) *cis*-AVP 28, and their crystal packing.

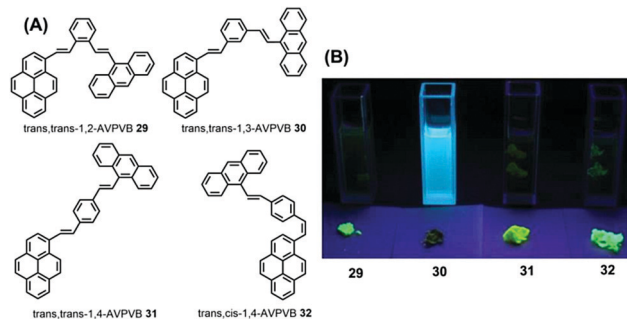


Fig. 19 Structures and fluorescence images of 29–32 in CH_2Cl_2 solution (upper) and as solids (lower) (from left to right) under illumination by a 365.0 nm UV lamp.

Due to differences in the molecular conformations, the series of compounds exhibit different fluorescence properties. For example, *trans,trans*-1-(9-anthryl)vinyl-4-(1-pyrenyl)-vinylbenzene (*trans,trans*-1,4-AVPVB 31) shows unusual AIE characteristics with fluorescence solvatochromism and emission shifts of over 80 nm, while *trans,trans*-1-(9-anthryl)vinyl-3-(1-pyrenyl)-vinylbenzene (*trans,trans*-1,3-AVPVB 30) exhibits a weak solvatochromic effect due to the *meta*-substitution which emits most efficiently in solution. In addition, *trans,trans*-1,2-AVPVB 29 and *trans,cis*-1,4-AVPVB 32 display moderate fluorescence and solvatochromism.

Theoretical calculations indicate that fluorescence solvatochromism is related to the conformational change from the ground state to the excited states due to the conjugate effect. An X-ray diffraction study revealed that the dihedral angle (53.98°) between the pyrenyl and central benzene ring in *trans,cis*-1,4-AVPVB 32 is higher than that observed for *trans,trans*-1,4-AVPVB 31 (22.46°), which is due to the strong intramolecular C-H $\cdots\pi$ steric hindrance between the pyrenyl ring and its adjacent phenyl in *trans,cis*-1,4-AVPVB 32. Therefore, the *trans,trans*-1,4-AVPVB 31 molecule is more planar and has better π -conjugation than does *trans,cis*-1,4-AVPVB 32.⁵³

On the other hand, α -cyanostyrenes of type 33 bearing a pyrene scaffold are weakly emissive in solution but show broad red-shifted emission in water due to aggregation (Fig. 20).⁵⁴

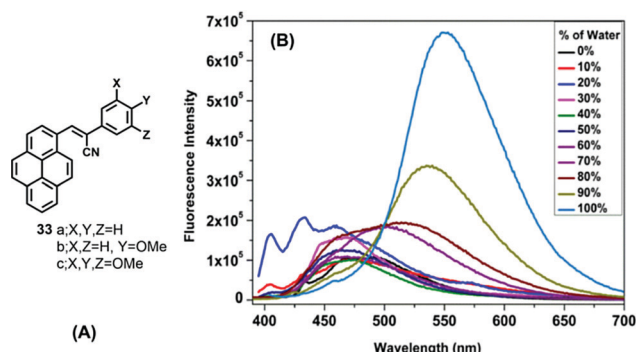


Fig. 20 (A) Structures of 33 and (B) emission of 33c in a 1,4 dioxane/water system. Reproduced from ref. 54.

Methoxy-substituted pyrenyl styrenes (**33b** and **33c**) exhibit a broad emission peak ($\sim 550\text{--}560\text{ nm}$) with a shoulder peak ($\sim 460\text{ nm}$) in polar acetonitrile. In the solid state, **33a** shows a clean emission spectrum with maximum peak at $\sim 510\text{ nm}$, while **33b** and **33c** exhibit a noisy emission spectrum with bands at $\sim 495\text{ nm}$ and at $\sim 530\text{--}540\text{ nm}$, respectively. The presence of the methoxy groups led to the formation of uniformly shaped aggregates, and the resulting unique emission behavior.⁵⁴ On the other hand, the emission of **33c** exhibited a bathochromic shift in different water fractions ($\lambda_{\text{em}} = 516\text{ nm}$ in $f_w = 60\%$, $\lambda_{\text{em}} = 535\text{ nm}$ in $f_w = 90\%$, $\lambda_{\text{em}} = 550\text{ nm}$ in $f_w = 100\%$), due to the formation of J-aggregates dominated by head-to-tail interactions.

2.4. Pyrene-based chalcone AIEgen

Pyrene-based chalcone derivatives, formed by combining chalcone units with a pyrene core, exhibit special optoelectronic properties, due to a push-pull conjugation effect that improves the luminescent properties and redox activity, while the α,β -unsaturated ketone linkage enhances the non-linearity.^{55,56} The molecular structure of such pyrenes consists of a donor with an electron-deficient keto vinyl (acceptor) linkage, which helps construct a D- π -A type conformation, which is responsible for the interesting AIEE/AIE features which enhance the luminescence properties of these materials. For instance, the four chalcone dyes **34**–**35** emitted yellow-green light in the solid state, and showed a large Stokes shifts *versus* their solution behavior.⁵⁵ The fluorescence spectra of the methoxy-containing dyes **34b** and **35b** are found to be blue-shifted by 10 nm relative to the parent ones (**34a** and **35a**), and a beneficial effect of the methoxy function on the degree of aggregation was observed. Quantum yields of **34b** (16%) and **35b** (19%) were found to be higher in the solid state than in solution, which thus showed a clear AIEE characteristic.

The X-ray crystal structure of **35b** revealed stacks of head-to-head-oriented chromophores (Fig. 21).⁵⁵ The case of the dye **35a** is different for two reasons; firstly, it forms dimers in which the π -systems of the pyrene styrenyl units in adjacent molecules are parallel, oriented head-to-head, and have no π -overlap.

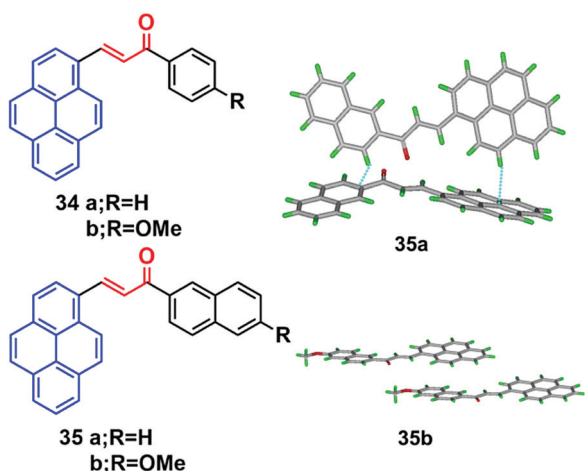


Fig. 21 Crystal packing diagrams of **34** and **35** illustrating π - π stacking interactions. Reproduced from ref. 55.

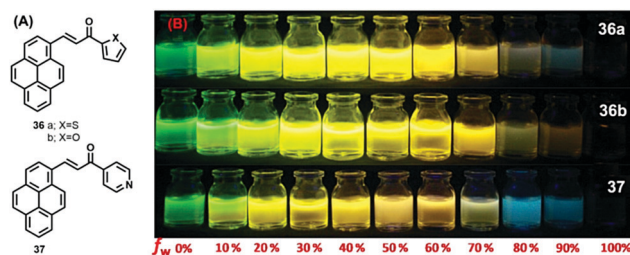


Fig. 22 Photograph on the AIEE properties of PCT **36a**, PCF **36b** and PCP **37** in THF/water mixtures with different water fractions ($f_w = 0\text{--}100\%$) under illumination by long UV-light (365 nm). Reproduced from ref. 56.

Secondly, the pyrene moieties of adjacent dimers of the former type overlap in an antiparallel fashion with an interplanar distance of $3.489\text{ \AA}$. The X-ray structure of **35a** reveals the presence of H-like aggregates, which may explain the quenching of the photoluminescence in these systems as compared with their CH_2Cl_2 solutions (Fig. 21). The X-ray structure of **35b** indicates a J-type aggregation that is not detrimental to solid-state emission.⁵⁵

On the other hand, pyrene-based chalcone materials containing heterocyclic rings (**36**–**37**) were synthesized by a Claisen–Schmidt condensation, and exhibit an AIEE with dual emissive characteristics in solution as well as in the solid.

Specifically, the compounds **36**–**37** showed a large tuneable color emission from green to orange, depending on the water fraction in the THF/ H_2O mixtures (Fig. 22).⁵⁶ In the solid state, three compounds underwent a red-shift of 77 nm , 82 nm , and 81 nm of **36a** (PCT), **36b** (PCF), and **37** (PCP) compared to their solution state, respectively. In the solid-state emission spectra, **36a** exhibits a maximum emission at 558 nm , whereas **36b** and **37** display maxima at 569 and 568 nm , respectively,⁵⁶ which is attributed to intermolecular interactions of the pyrene unit (namely π - π stacking and a J-aggregation arrangement). Indeed, compound **36a** (PCT) afforded a good emission quantum yield of 0.74 , whereas for **36b** (PCF) and **37** (PCP), the quantum yields were 0.68 and 0.53 , respectively.⁵⁶

2.5. Pyrene-based AIE-active Schiff base molecules

Many scientists are attempting to use the AIE activity of pyrene derivatives to sense molecules by using push-pull or on-off procedures. Schiff base AIE systems have been identified on the basis of intensity turn-on at a specific wavelength.⁵⁷ Novel pyrene-based Schiff base derivative **38** was synthesized *via* a condensation reaction and utilized as a turn-on fluorescence sensor toward Cu^{2+} , and showed AIE activity in CH_3CN and water mixture up to 80% (Fig. 23).⁵⁸ The PL intensity of **38** was enhanced *ca.* 630 -fold ($\Phi_f = 0.567$ in $\text{CH}_3\text{CN}:\text{H}_2\text{O} = 20:80$) by increasing the water fraction of H_2O from 0% to 90% and with a red-shift from 455 to 471 nm . The mechanism possibly arises from RIR. Since the single-bond rotation is mainly responsible for the dominant non-radiative decay, the RIR effect might be the cause of the AIE nature of **38**. The hydrogen-bond formation, suppression of the photo-induced electron transfer (PET) process, and suppression of charge transfer (CT) or even intramolecular

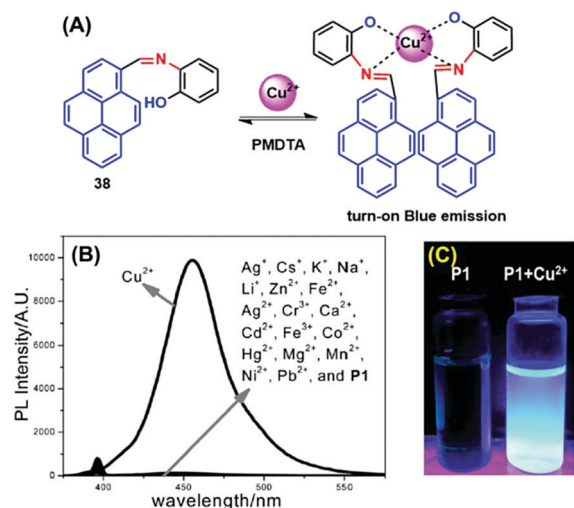


Fig. 23 (A) Possible proposed binding mechanism of **38** toward Cu^{2+} ions. (B) Fluorescence emission spectra of compound **38** (10 μM) upon the addition of various metal ions. (C) Photographs of **38** and **38** + Cu^{2+} visualized under UV-light irradiation. Reproduced from ref. 58.

charge transfer (ICT) are other mechanistic approaches for this category of AIEgens.⁵⁸ However, the PL spectra for **38** suggested that the red-shift on aggregation probably originated from the suppression of the TICT. Various molecules act as turn-on sensors for a variety of cations and anions based on PET, intramolecular charge transfer (ICT), chelation-enhanced fluorescence (CHEF) and deprotonation mechanisms.⁵⁹

The photophysical properties of **39** revealed that the fluorescence of this molecule is completely quenched due to PET processes in solution. The fluorescence intensity remained unchanged over the water fraction range 0–70%. On further addition of water up to 90%, the emission intensity was enhanced with a red-shift (25 nm) from 440 to 465 nm, with a concomitant increase in fluorescence lifetime.⁶⁰ This clearly signifies that this molecule exhibited an AIE property (Fig. 24).

The AIE behavior of **39** in the excited state was further investigated by the time-correlated single-photon counting (TCSPC) technique, and the observed longer fluorescence lifetime is due to suppression of the PET process caused by aggregation of the

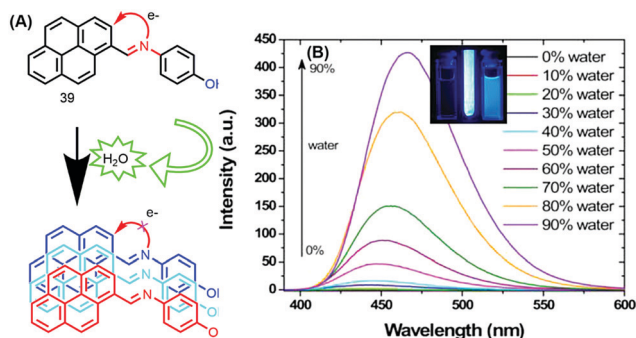


Fig. 24 (A) **39** in THF and (B) emission spectra of **39** in THF with increasing percentages of water; inset: photographs taken under illumination by a UV lamp in a THF/water mixture. Reproduced from ref. 60.

molecules of **39**.⁶⁰ Thus, the mechanism of AIE for this molecule is suppression of PET due to the hydrogen-bonding interactions of the imine donor with water.

Tunable pyrene-based benzohydrazonate molecules of type **40** exhibit drastic aggregation-induced color changes ($\Delta\lambda_{\text{em,max}} \sim 130$ nm) with high solid-state quantum yields ($\sim 54.5\%$). The exciting switch behavior is triggered by intermolecular hydrogen bonding *via* the carbohydrazone bridge, which leads to restriction of the molecular rotation (Fig. 25).⁶¹ Upon increasing the amount of water in the dimethyl formamide (DMF) solution, the absorption did not significantly change until the water fraction reach 90% for ($\text{R} = \text{H}$) and 80% for ($\text{R} = \text{Cl}$, OMe), whereupon aggregation was initiated. The solid compounds **40** ($\text{R} = \text{Cl}$, OMe) revealed red-shifts of 54 nm and 42 nm compared with the spectra of the monomer state.

Furthermore, compound **40** exhibited mechanochromic properties upon exogenic action. The blue-emitting pristine solids were converted to yellowish-green emitting solids, and the emission peaks exhibited varying degrees of red-shift. The pyrenyl-substituted acylhydrazone **41** derivative also exhibited favorable multiple luminescent switching mechano-fluorochromic behavior. In particular, fluorescence color changes were visible by the naked eye, as well as emission that could be reversibly switched between a blue and a green color, as shown in Fig. 26.⁶² By contrast, xerogel **41** obtained from dimethyl sulfoxide (DMSO) revealed mechanofluorochromic behavior, with the maximum emission peak shifted from 458 nm to 514 nm upon grinding. However, for **41** precipitated from THF, no mechanofluorochromic behavior was observed. The different luminescent properties of **41**

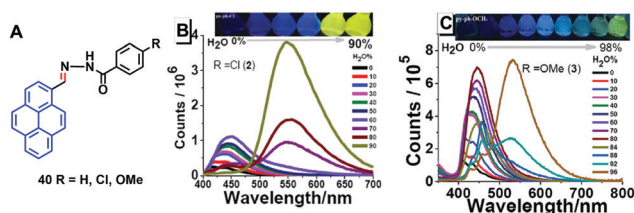


Fig. 25 Structures and PL spectra of **40** in DMF/ H_2O mixtures with different volume fractions of H_2O . Inset: Emission color change from blue to yellow in aqueous DMF with $f_w = 0$ –90% under 365 nm UV-light. Reproduced from ref. 61.

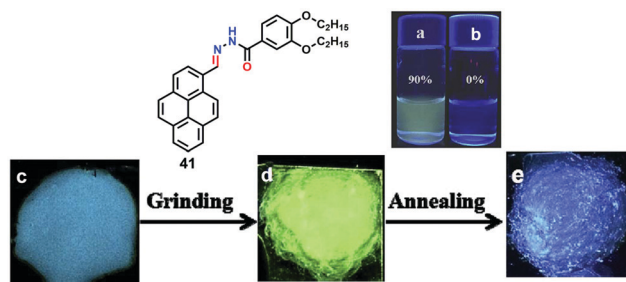


Fig. 26 Emission spectra of **41** in 90% water (a) and THF (b). Photographs of emitting samples upon 365 nm UV illumination: (c) xerogel **41**, (d) ground xerogel and (e) after annealing treatment. Reproduced from ref. 62.

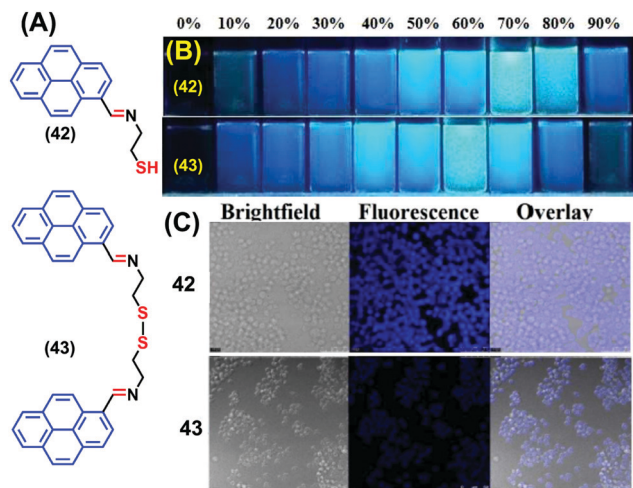


Fig. 27 (A) Structures of **42** and **43**. (B) Photograph of the AIE of **42** and **43** visualized under UV-irradiation ($\lambda = 365$ nm). (C) Fluorescence images of Raw264.7 cells treated with **42** and **43** at 12 h. Reproduced from ref. 63.

are attributed to the switch of the self-assembled structures *via* intermolecular hydrogen bonding which impacted on the translational and rotational motion of the molecules, and thus the conformational flexibility of the pyrenyl moieties was restricted.⁶²

Two novel pyrene-containing monomeric and dimeric Schiff base derivatives **42** and **43** were reported to show AIEE activity in DMSO upon addition of H₂O over the range 0–90% (Fig. 27).⁶³ Both **42** and **43** were also good sensors for trivalent metal ions (Fe³⁺, Cr³⁺ and Al³⁺) in living cells. As an example, when Raw264.7 cells were incubated with **42/43** (10 mM in DMSO), no fluorescence was observed over 30 min. However, bright blue fluorescence images were observed for the Raw264.7 cells after 12 h with **42** and dismal blue fluorescence images with **43** due to intracellular H₂O-induced aggregation. The bright cell image of **42** at 12 h is due to the greater intracellular penetration of the free thiol unit, which is superior to the disulfide-containing **43**.⁶³ Both **42** and **43** showed “off-on” sensing at highly acidic pH values (1–3) and were employed for live cell applications.⁶³

2.6. ESIPT-based AIEgens

ESIPT compounds have attracted attention due to their interesting and unique emission properties. Essentially, introducing the AIE phenomenon and excimer formation ability into molecular systems has provided the scope for dual emission. By contrast, ESIPT-coupled AIE phenomena operate as an additional handle to adjust the emission wavelength of the corresponding emission peak by varying the solvent polarity.

The development of ESIPT molecules as probes is rapidly growing given the recent interest in their photophysical properties, such as unusually large Stokes shifts and outstanding environment-sensitivity that can readily be monitored by spectral changes.⁶⁴ It should be noted that ESIPT molecules exist exclusively in an enolic form (E) in the ground state. Upon photoexcitation, extremely fast tautomerism from the excited enolic (E*) to the excited keto (K*) form occurs on a sub-picosecond time scale. After decay of K* to the ground state, the K form transforms to the initial E form through

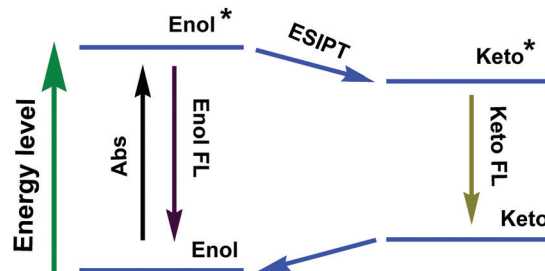


Fig. 28 Four-level schematic diagram of ESIPT.

reverse proton transfer mediated by intramolecular hydrogen bonds (Fig. 28). The different species between absorbing and emitting in the intrinsic four-level photocycle (E–E*–K*–K) offer a high probability of addressing the disadvantage of fluorescence concentration quenching that is suffered by almost all of the currently available luminescent probes. On the other hand, ESIPT involves a proton-transfer process that is easily affected by changes in the surrounding conditions, leading to sensitive spectral responses. A single and broad emission peak due to complete ESIPT is observed in polar solvents, while in non-polar solvents, conversion of the enol into the keto form is incomplete. The J-type aggregate effects may block the molecular motion and facilitate the ESIPT emission in the solid state/aggregated state.

Samanta *et al.* designed a simple AIE-ESIPT-coupled organic molecule bearing a suitable fluorophore for achieving tuneable white-light emission.⁶⁵ The designed molecule **44** was found to be very weakly emissive in pure methanol, but this was enhanced in methanol–water mixtures with large water contents (Fig. 29).

A pyrene-based fluorescent probe, 3-methoxy-2-((pyren-2ylimino)methyl)phenol **45** (Fig. 30), was found to exhibit high sensitivity and selectivity toward Al³⁺ over other relevant metal ions, and also displayed novel AIEE characteristics in its aggregated/solid state.⁶⁶ This material could sense Al³⁺ in a 1:1 mode by signaling a significant fluorescence enhancement with a turn-on ratio of over ~200-fold. Excess addition of Al³⁺ dramatically enhanced the fluorescence intensity through aggregation. The origin of such an AIEE phenomena for **45** was due to the single-bond rotation, which is mainly responsible for the dominant non-radiative decay.⁶⁶ Thus, the RIR of **45** plays a crucial role in the

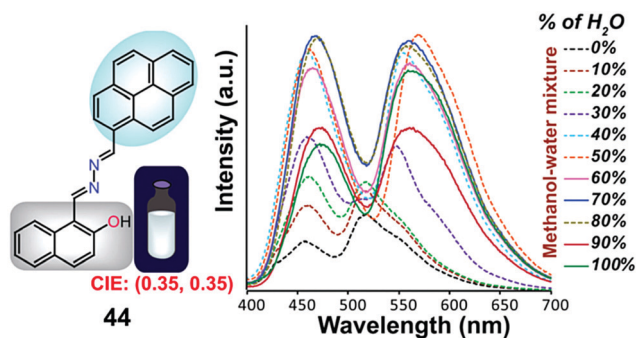
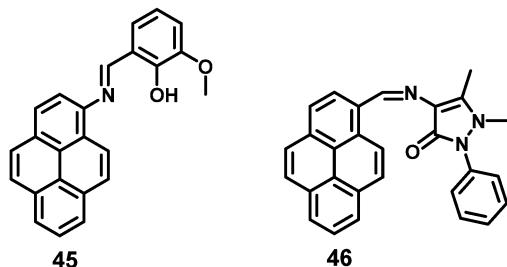


Fig. 29 Structure of **44** and fluorescence spectra (10 mM) upon a change in the water fraction of a methanol–water mixed solvent; $\lambda_{\text{ex}} = 380$ nm. Reproduced from ref. 65.

Fig. 30 Structures of **45** and **46**.

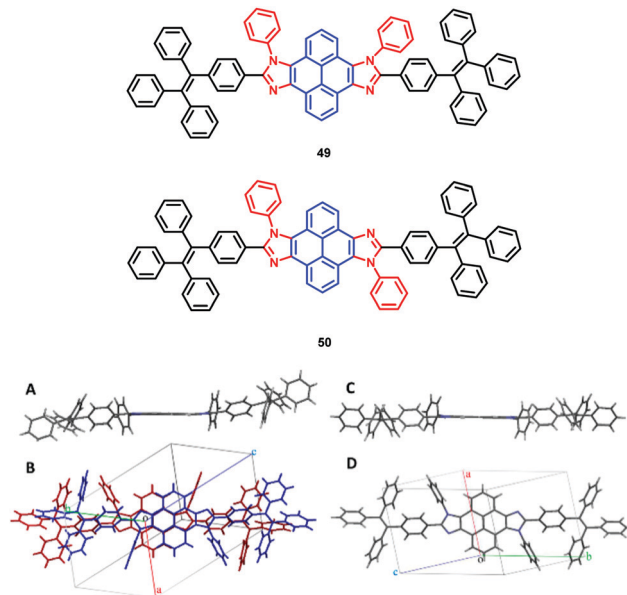
AIEE characteristics as well as the PET off in the solid state. The same group also reported the AIEE materials **46** (Fig. 30), which can sense 2,4,6-trinitrophenol (TNP) by PET off and exhibit AIE in the aggregation state.⁶⁷

2.7. Non-typical pyrene-based AIEgens: modification at the K-region

The pyrenoimidazoles **47** and **48** are weakly fluorescent in solution, and their AIE behavior has been studied in THF–water mixtures with different water content.⁶⁸ The poor fluorescence of the pyrenoimidazoles **47** and **48** in THF was enhanced in the aggregated suspension (95% water fraction) by 38- and 9-fold, respectively (Fig. 31). The pristine forms of pyrenoimidazole absorb at 417 nm and 427 nm, respectively, which upon grinding, exhibit bathochromic shifts and absorb at 446 nm and 437 nm, respectively. RIR and X-ray diffraction (XRD) studies reveal a transformation of the twisted crystalline state to the planar amorphous state.⁶⁸

The pyrene-imidazole-based structural isomers, namely axis symmetric *syn*-PyDTI **49** and centrosymmetric *anti*-PyDTI **50**, are non-emissive in THF but exhibited intense emission in their aggregate states. Crystalline *syn*-PyDTI **49** exhibited a 2-fold higher quantum efficiency than *anti*-PyDTI **50**, and displayed a much better performance in OLEDs with a $\eta_{\text{C,max}}$ of 11.4 cd A^{-1} (8.8 cd A^{-1} for **50**).⁶⁹

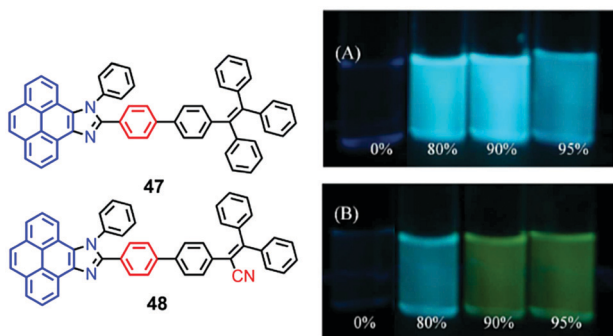
Comparing the crystal emissions revealed a red-shift for the *anti*-isomer **50**; that is, the *syn*-PyDTI **49** crystal exhibited an emission peak at 465 nm, whereas the *anti*-isomer **50** crystal displayed an emission peak at 507 nm. High-quality single

Fig. 32 (A and C) Structures of *syn*-PyDTI **49** and *anti*-PyDTI **50**. (B and D) Unit cell structure top views of crystalline *syn*- and *anti*-PyDTI. Reproduced from ref. 69.

crystals of *syn*- and *anti*-PyDTI emit blue and green, respectively, under UV-light. The blue-shift observed for *syn*-PyDTI **49** was probably associated with the formation of dimers, which resulted in a flexible geometry and a decreased conjugation length (Fig. 32).⁶⁹ These observations reveal the dependence of the molecular excited-state properties on distinct structural symmetries. In recent years, pyrene-based, mechanochromism (MC)-active materials have been reported to possess attractive excimer fluorescence in the solid state, which is in contrast to the frequently encountered ACQ phenomena in planar luminogens and offers the possibility to discover mechanoluminescence (ML) luminogens with planar structures.^{70,71}

An outstanding feature associated with pyrene is its facile and varied structural modification. In addition to its four main points of reactivity (*i.e.*, the 1-, 3-, 6- and 8-positions), recent studies have revealed that the non-typical K-region (4-, 5-, 9-, and 10-positions) is also available for further modification.^{2,9,72} In contrast to these recent advances at the 1-, 3-, 6- and 8-positions, pyrene derivatives substituted at the 4-, 5-, 9-, and 10-positions (K-region) have been employed predominantly as synthetic intermediates for fused π -systems.⁷³ However, only a few studies have examined the effects of substitution at the 4-, 5-, 9-, and 10-positions on the photophysical behavior of pyrenes.⁷⁴

A new design strategy was introduced by the Sasaki group in order to achieve AIE emission in pyrene chromophores. This was accomplished by introducing *N,N*-dimethylamino substituents at the K-region 4,5-positions of pyrene, and these are likely to stabilize the minimum energy conical intersection (MECI).⁷⁵ Four pyrene derivatives, which contain highly twisted *N,N*-dimethylamino groups at the 4- (4-Py), 4,5- (4,5-Py), 1- (1-Py), or 1,6-positions (1,6-Py) are **51–54** (Fig. 33). The non-radiative transitions of 4,5-Py **52** are highly efficient ($k_{\text{nr}} = 57.1 \times 10^7 \text{ s}^{-1}$), so that its fluorescence quantum yield in acetonitrile decreases to $\Phi = 0.04$.⁷⁵ The solid-state fluorescence of 4,5-Py **52** is also

Fig. 31 Photographs of **47** (upper) (A) and **48** (lower) (B) in THF–water mixtures with different water fractions (10 mM) under 365 nm UV illumination. Reproduced from ref. 68.

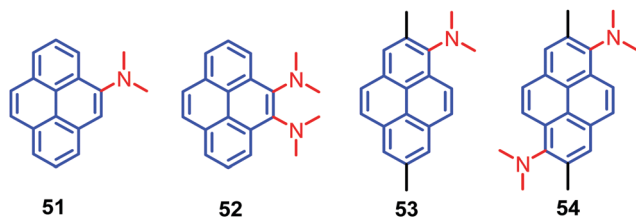


Fig. 33 Structures of 51–54.

efficient ($\Phi = 0.49$). By contrast, 1,6-Py 54 features strong fluorescence ($\Phi = 0.48$) in solution with a slow non-radiative transition ($k_{nr} = 11.0 \times 10^7 \text{ s}^{-1}$) that is subject to severe quenching ($\Phi = 0.03$) in the solid state. The introduction of strongly twisted *N,N*-dimethylamines at the 1- and 1,6-positions resulted in intense fluorescence in solution. These differences are attributed mainly to non-radiative transitions. A comparison of the k_{nr} value of 4,5-Py 52 with its fluorescence energetics implies the involvement of a special pathway; *e.g.* a non-adiabatic relaxation *via* MECI, for efficient internal conversions.⁷⁵ The other derivatives 51 and 53 did not show noticeable emission changes in solution or the solid state, although they have AIE activity. The presence of highly twisted *N,N*-dialkylamino groups, especially at the 4,5-positions, induces fast non-radiative transitions both in polar and non-polar solvents, resulting in faint fluorescence in solution.⁷⁵

Interestingly, a pyridine unit was introduced at the 1-position of pyrene, affording 2-(pyren-1-yl)pyridine 55 (Fig. 34).⁷⁶ This compound exhibits a typical AIEE feature, and mechanisms were proposed by RIR, while the crystal structure of 55 revealed the presence of a C–H... π interaction at 2.75 Å between two of the pyridine rings of adjacent molecules which blocks the rotation of the pyridine ring in the solid state (Fig. 34). Furthermore, strong π – π interactions were observed in the range of 3.53–3.59 Å. Such short contacts would suppress the restriction of movement of the pyridine group in the solid state, and block

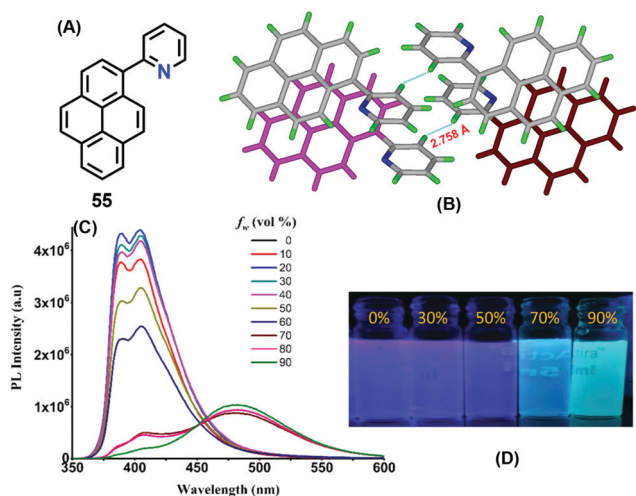


Fig. 34 Emission spectra of 55 in different fractions of water in methanol and strong π – π interactions between the centroids of the two pyrene rings in the solid state. Reproduced from ref. 76.

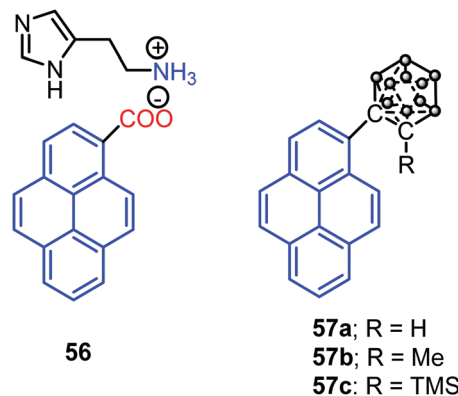


Fig. 35 Structures of 56–57.

the non-radiative pathways and open up new radiative pathways on the progressive addition of water. On the other hand, the face-to-face π – π interaction also caused an excimer emission in the aggregation state or in the crystal state.

Interestingly, Dastidar *et al.* first observed the AIE phenomenon in organic salts. The primary ammonium salts of the pyrene-based carboxylic acid and histamine 56 (Fig. 35) exhibited an enhanced fluorescence in the solid state. The N acceptor and N–H donor of the imidazole moiety form a 3D hydrogen bonding network, thereby imposed RIR in the solid state, and triggering AIE activity in the pyrene salt 56.⁷⁷ The pyrene-substituted *o*-carborane dyads 57 (Fig. 35) were synthesized *via* the insertion reaction between decaborane and 1-ethynylpyrene in the presence of a Lewis base using the Sonogashira–Hagihara coupling reaction. *O*-Carborane dyads emit extremely bright light in the solid state ($\Phi_{PL} > 0.99$).⁷⁸

2.8. Pyrene-based hexaphenylbenzene-like AIEgens

As mentioned above, steric interactions play an important role and can affect molecular configuration, as well as luminescence behavior. For example, the propeller-shaped hexaphenylbenzene (HPB), consists of six rotatable peripheral phenyl units with respect to the central benzene ring, and shows AEE character.⁷⁹ With this in mind, new pyrene-based hexaphenylbenzene-like AIEgens 58 were explored, as shown in Fig. 36. Two phenyl units were replaced by a pyrene ring, and both *anti*-58 and *syn*-58 presented novel AIEE characteristics with a multiple photoluminescence, including monomer, excimer and CT emission. The multiple emission mechanisms originated from the different spatial arrangements, due to the presence of the large planar structure of the pyrene unit.⁸⁰

3. Possible applications of pyrene-based AIEgens

“Touch a stone and turn it into gold”: pyrene, a dim luminogen in solution, can be twisted into a bright emitter as a solid by decorating it with suitable groups at different active positions. Solid-state fluorescent organic materials are essential components in optoelectronic devices, such as OLEDs, solid-state lasers fluorescent, sensors, organic solar cells (OSCs).⁸ However, the emission of most

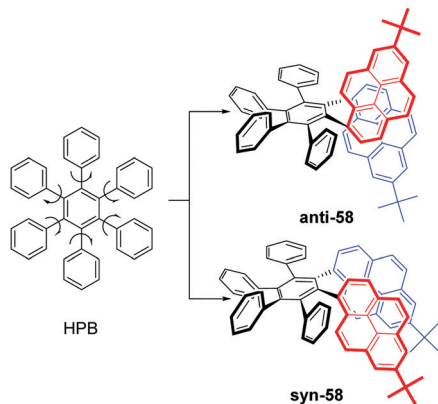


Fig. 36 Structure of HPB and the pyrene-based hexaphenylbenzene-like AIEgens **58**.

fluorescent organic molecules is effectively suppressed in the solid state, due to aggregation of the planar π -conjugation systems based on π - π interactions or the formation of excimers.⁷ As a result, there has been considerable interest in preparing pyrene-based AIEgens which retain their photophysical properties in the solid state. Pyrene-containing probes are very topical and remain of widespread interest because of their fast response *via* PET and excimer/exciple formation toward specific species of interest.⁸¹

3.1. Organic light-emitting diodes

Pyrene is well known as a good deep blue emitter and is often utilized to construct blue luminogens.^{22a,82} An OLED, also termed organic electro luminescent device (OELD), is a light-emitting diode (LED) whose emissive electroluminescent layer is composed of an organic film. This layer of organic semiconductor material is formed between two electrodes, where at least one of the electrodes is transparent. An OLED material as a flat-panel display has many unique advantages, such as low driving voltage, high efficiency, excellent flexibility, vivid color, wide viewing angle. As a consequence, this technology has been used successfully for solid-state lighting applications. Pyrene is a promising chromophore with a high fluorescent quantum yield, and has attracted considerable interest in the design of small organic molecules for different applications including OLEDs⁸³ and organic field-effect transistors (OFETs).^{70,84} For example, the optimization of non-doped OLEDs using pyrene-based 2,7-TPE-substituted AIEgens led to emitters that displayed sky-blue light at 492 nm at a low turn-on voltage (3.1 V) with maximum luminance of 15750 cd m⁻², current efficiency of 7.34 cd A⁻¹, and with low efficiency roll-off (Fig. 37).⁸⁵

N,N-Di-*p*-methoxyphenylamine-substituted pyrenes have been successfully used as hole-transporting materials for perovskite-based solar cells with high efficiency.⁸⁶ Since pyrene has extensive π -electron delocalization, it is a fascinating core for developing fluorescent π -conjugated light-emitting monomers substituted by a variety of functionalized groups or can be deployed themselves as substituents and introduced into others molecules to afford blue EL.

Advantages of pyrene **1** include its excellent thermal and chemical stability, but it is not a good candidate for LEDs alone,

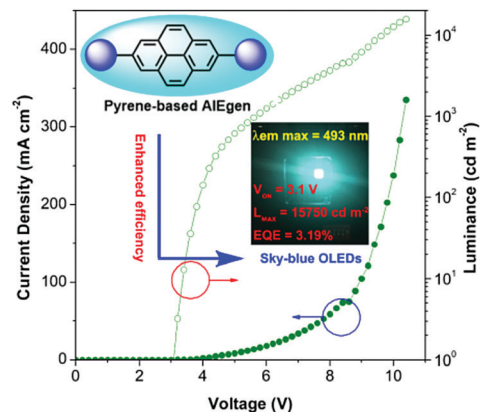


Fig. 37 Using 2,7-disubstituted pyrenes with AIE characteristics as an emitter, the sky-blue OLED device exhibited excellent EL properties. Reproduced from ref. 85.

due to its excimer emission tendency which lowers the fluorescence efficiency.¹⁰ However, this limitation has been overcome by introducing bulky groups with long alkyl chains, which enhances the fluorescence quantum efficiency in OLEDs and also reduces the aggregation of the final material.²

3.2. Chemosensors

As a classic chromophore, pyrene shows interesting fluorescence properties and has been used in chemosensors.^{23a,87} Pyrene-based AIE/AIEE probe molecules behave as selective and sensitive detectors of metal ions (Cu²⁺, Fe³⁺, Al³⁺), picric acid, and H⁺ ions *etc.*^{58,66,67,78}

3.3. Organic field-effect transistors

The high fluorescence efficiency, electron rich 'macrocyclic' conjugation, and high charge carrier mobility all make pyrene a promising candidate for a variety of organic semiconducting materials.⁸⁸ Pyrene-based organic semiconductors have great opportunities in OFETs as charge carrier channel layer materials possessing transport mobility.^{8b,87b,89}

Interestingly, Zhang *et al.* have reported two *trans*- β -styryl substituted pyrene-based AIEgens 1,6-PyE **59** and 2,7-PyE **60** which exhibited strong emission with semiconducting properties in the solid state.⁹⁰ Microrods of 1,6-PyE **59** exhibit intense yellow-green emission ($\lambda_{\text{max,em}}$: 463 nm), while microplates of 2,7-PyE **60** show brighter blue-emissions ($\lambda_{\text{max,em}}$: 437 nm) with high quantum yields of 28.8% and 27.4%, respectively. Furthermore, thin films of the pyrene-based AIEgens, **59** and **60** (Fig. 38) exhibited p-type semiconducting properties with hole mobility up to 1.66 cm² V⁻¹ s⁻¹, which originate from their unique intermolecular interactions and crystal packing. In addition, both compounds exhibited optical waveguiding behavior.

3.4. Solid-state lasers

Recently, a considerable amount of effort has been devoted to the design of laser materials and devices because of their potential applications.⁹¹ Solid-state lasers comprise a host and an active ion doped host material. The active ion should have excellent fluorescent, broad absorption bands, high quantum yield *etc.*,

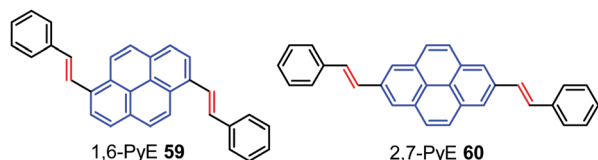


Fig. 38 Molecular structures of 1,6-PyE **59** and 2,7-PyE **60**.

and since pyrene possesses these properties, it is a good candidate for the design and development of new solid-state lasers.

3.5. Biological applications

Pyrene-based AIE compounds show good biological activity against both Gram-positive (*Rhodococcus rhodochrous* and *Staphylococcus aureus*) and Gram-negative (*Escherichia coli* and *Pseudomonas aeruginosa*) bacteria and also fungal species (*Candida albicans*).⁶⁰ Importantly, AIEE probes can sense M^{3+} ions, and their highly acidic pH-sensing ability has been successfully applied in cell imaging and cell viability analysis.⁶³ The sensitive and selective detection of nucleic acids is important in biological studies, clinical diagnostics, and bio-defense applications, and pyrene-based AIEgens have potential in these fields.⁹²

3.6. Probing critical micellar concentration

The critical micellar concentration (cmc) of both ionic and non-ionic surfactants can conveniently be determined from measurement of the UV absorption of pyrene in surfactant solution. Surfactants can assemble in solution, and cmc is an important solution property of surfactants. There are several frequently used methods such as tensiometry, conductometry, fluorimetry, and calorimetry, which can be used for its determination.⁹³ Spectral methods using dyes as probes, and self-absorption of amphiphiles in solution are used for the evaluation of cmc.⁹⁴ As fluorescent probes, pyrene molecules have been successfully employed to estimate the polarity level of micelles as well as the cmc of amphiphiles.⁹⁵

3.7. Mechanoluminescent and mechanochromism agents

The solid-state behavior of organic materials, which exhibit bright emission under mechanical actions such as pressure, grinding, shearing, rubbing, temperature *etc.*, is attracting much attention due to its scientific significance and practical applications in many fields, including uses in deformation detectors, new light sources, and pressure sensing.⁹⁶

For example, Zhen Li *et al.* reported the unique solid-state packing style of pyrene and the bright dual monomer–excimer-ML and excimer-ML from two pyrene derivatives (Py-Bpin **61b** and Py-Br **61c**). The simple, planar molecular structure was compared with pyrene and the analog Py-H **61a** (Fig. 39).⁹⁷ These four luminogens display similar optical properties in dilute solution, but different luminescent properties in the solid state, mainly due to their different molecular packing in the crystalline form. For example, pyrene is MC- and ML-inactive, while Py-H **61a** is MC-active, Py-Bpin **61b** is MC- and ML-active, and Py-Br **61c** is ML-active. The MC-active pyrene derivatives (Py-H **61a** & Py-Bpin **61b**) have monomer arrangements with very

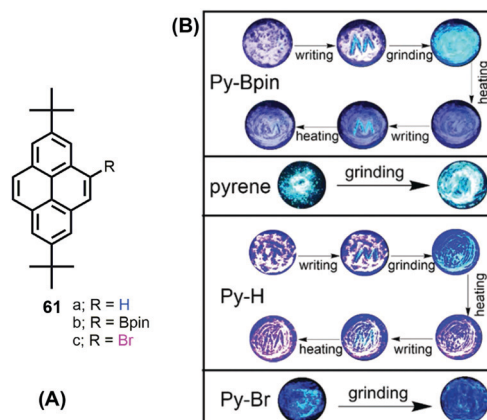


Fig. 39 Structures and images of mechanochromic experiments involving pyrene **1**, Py-Bpin (**61b**), Py-H (**61a**) and Py-Br (**61c**). Reproduced from ref. 97.

few overlaps. The ML-active pyrene derivatives (Py-Bpin **61b** & Py-Br **61c**) have strong intermolecular interactions in the solid state. Both a monomer arrangement and strong intermolecular interactions are required in MCL-active pyrene derivatives (Py-Bpin **61b**). These results demonstrate that planar molecules can also have bright ML if the luminogens possess strong intermolecular interactions and no ACQ in the solid state.

4. Perspectives and conclusions

In conclusion, pyrene is an important member of the PHA family, which has received considerable attention in a variety of applications, and recent reports in the literature reveal that it continues to attract the interest of chemists and those involved in materials science and beyond. To further expand the pyrene chemistry research, the basis of this review is to focus on summarizing the efficient strategies used to develop pyrene-based AIEgens, which are very efficient and promising materials with fantastic optical properties. This includes (1) it is possible to achieve pyrene-based AIEgens by introducing typical AIE units, such as TPE, TriPE, into the active site of pyrenes. (2) It seems that C=C or C=N bonds play an important role in constructing the pyrene-based AIEgens and this molecular configuration can contribute to the molecular rotation against the C=C/C=N bond around the single-bond axes which non-radiatively consume the exciton energy. (3) The molecular (crystal) packing can affect the photophysical properties, and can lead to interesting optical phenomena, such as mechanoluminescent/mechanochromism. Thus, the twisted molecular structure bearing a bulky group can suppress intramolecular rotation and enhance the intensity of fluorescence in the aggregated state. (4) In addition, the pyrene-based primary ammonium salts also exhibited an enhanced fluorescence in the solid state, which may be due to the anion- π^+ interactions suppressing any π - π stacking. This is useful approach for the design and synthesis of pyrene-based AIEgens.⁹⁸

The electronic composition at the active sites (1,3,6,8-positions), the nodal plane (2,7-positions), and the K-region

(4,5,9,10-positions) of pyrene are all different. When terminal groups are attached at the 1,3,6,8-positions, this leads to a large red-shift in the emission compared to when they are positioned at the nodal plane and the K-region.

By contrast, substitution at the 2,7-positions of pyrene has only a limited effect on the emission shift, due to the weak electronic interaction communication between the pyrene core and the substituted moieties. The K-region of pyrene exhibits a more special electronic structure, the substituent groups at the 4,5,9,10-positions can affect the $S_1 \leftarrow S_0$, $S_2 \leftarrow S_0$ and $S_3 \leftarrow S_0$ transitions.⁹⁹ According to the Jablonski diagram, 25% of the singlet excitons can be utilized to produce light, while 75% of the radiative decay of triplet excitons is forbidden.¹⁰⁰ To improve the quantum yield of luminogens devices, there is needs to harvest both singlet and triplet excitons under electrical excitation for emission. There are two main strategies (i) convert the triplet excitons (75%) into singlet excitons *via* reverse intersystem crossing (RISC) processes under thermal activation¹⁰¹ and (ii) develop phosphorescent emitters with high efficiency.¹⁰² Thus, to achieve highly efficient AIEgens, it is better to design and synthesize novel pyrene-based AIEgens to overcome the ACQ effect. Moreover, the new concepts for molecular design, such as aggregation-induced delayed fluorescence (AIDF),¹⁰³ thermally activated delayed fluorescence (TADF),¹⁰⁴ and room-temperature phosphorimetry (RTP)¹⁰⁵ can be introduced into the pyrene-based AIE system to improve the efficient exciton for organic semiconductor applications.

This review will be helpful in understanding and designing new high-efficiency pyrene-based solid emitters for further investigations of different applications, such as OLEDs, biosensors, chemosensors, and chiral recognition. Such studies will also shed further light on the key roles of the AIE or AIEE processes in pyrene chemistry.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the National Natural Science Foundation of China (21602014), and the Youth Fund of Guangdong University of Technology (220413205). MM Islam thanks MOST, PR China for a Talented Young Scientist Program (TYSP) Fellowship; CR thanks the EPSRC for an Overseas Travel Grant.

Notes and references

- (a) S. Tao, Z. Peng, X. Zhang, P. Wang, C. S. Lee and S. T. Lee, *Adv. Funct. Mater.*, 2005, **15**, 1716; (b) J. Jayabharathi, P. Jeeva, V. Thanikachalam and S. Panimozhi, *J. Photochem. Photobiol. A*, 2017, **346**, 296; (c) J. K. Salunke, F. L. Wong, K. Feron, S. Manzhos, M. F. Lo, D. Shinde, A. Patil, C. S. Lee, V. A. L. Roy, P. Sonar and P. P. Wadgaonkar, *J. Mater. Chem. C*, 2016, **4**, 1009; (d) D. Chercka, S.-J. Yoo, M. Baumgarten, J.-J. Kim and K. Müllen, *J. Mater. Chem. C*, 2014, **2**, 9083.
- M. T. Figueira-Duarte and K. Müllen, *Chem. Rev.*, 2011, **111**, 7260.
- S. Karupppannan and J. C. Chambron, *Chem. – Asian J.*, 2011, **6**, 964.
- H. Maeda, T. Maeda, K. Mizuno, K. Fujimoto, H. Shimizu and M. Inouye, *Chem. – Eur. J.*, 2006, **12**, 824.
- Z. Wang, C. Xu, X. Wang, X. Dong, B. Zhao and B. Ji, *Dyes Pigm.*, 2011, **92**, 732.
- T. Oyamada and H. Uchiuzou, *J. Appl. Phys.*, 2005, **98**, 74506.
- (a) Z. Zhang, B. Xu, J. Su, L. Shen, Y. Xie and H. Tian, *Angew. Chem., Int. Ed.*, 2011, **50**, 11654; (b) Q. Wang, H. Lu, L. Gai, W. Chen, G. Lai and Z. Li, *Dyes Pigm.*, 2012, **94**, 183; (c) M. Shimizu and T. Hiyama, *Chem. – Asian J.*, 2010, **5**, 1516.
- (a) *Organic light-emitting materials and devices*, ed. Z. Li and H. Meng, CRC Press/Taylor & Francis, Boca Raton, 2007; (b) F. Cicoira and C. Santato, *Adv. Funct. Mater.*, 2007, **17**, 3421; (c) *Introduction to organic electronic and optoelectronic materials and devices*, ed. S. S. Sun and L. R. Dalton, Taylor & Francis, Boca Raton, 2008.
- (a) X. Feng, J.-Y. Hu, C. Redshaw and T. Yamato, *Chem. – Eur. J.*, 2016, **22**, 11898; (b) Y. Gong, K. Chang, C. Chen, M. Han, X. Zhan, J. Min, X. Jiao, Q. Li and Z. Li, *Mater. Chem. Front.*, 2019, **3**, 93.
- (a) J. B. Birks, *Photophysics of Aromatic Molecules*, Wiley-Interscience, London, 1970; (b) J. R. Lackowicz, *Principles of Fluorescence Spectroscopy*, Kluwer Academic/Plenum Publishers, New York, 2nd edn, 1999, p. 595.
- P. Sonar, M. S. Soh, Y. H. Cheng, J. T. Henssler and A. Sellinger, *Org. Lett.*, 2010, **12**, 3292.
- (a) M. J. J. Dewar and R. D. Dennington II, *J. Am. Chem. Soc.*, 1989, **111**, 3804; (b) H. Cerfontain, K. Laali and H. J. A. Lambrechts, *Recl. Trav. Chim. Pays-Bas*, 1983, **102**, 210.
- (a) X. Luo, J. Li, C. Li, L. Heng, Y. Q. Dong, Z. Liu, Z. Bo and B. Z. Tang, *Adv. Mater.*, 2011, **23**, 3261; (b) H. Tong, Y. Hong, Y. Dong, Y. Ren, M. Haussler, J. W. Y. Lam, K. S. Wong and B. Z. Tang, *J. Phys. Chem. B*, 2007, **111**, 2000.
- G. Venkataramana and S. Sankararaman, *Eur. J. Org. Chem.*, 2005, 4162.
- (a) C. Lambert, J. Ehbets, D. Rausch and M. Steeger, *J. Org. Chem.*, 2012, **77**, 6147; (b) L. Ji, R. M. Edkins, A. Lorbach, I. Krummenacher, C. Brückner, A. Eichhorn, H. Braunschweig, B. Engels, P. J. Low and T. B. Marder, *J. Am. Chem. Soc.*, 2015, **137**, 6750; (c) A. G. Crawford, A. D. Dwyer, Z. Liu, A. Steffen, A. Beeby, L.-O. Pålsson, D. J. Tozer and T. B. Marder, *J. Am. Chem. Soc.*, 2011, **133**, 13349.
- C. Wanninger-Weiß and H.-A. Wagenknecht, *Eur. J. Org. Chem.*, 2008, 64.
- (a) J. Wang, Y. Zhao, C. Dou, H. Sun, P. Xu, K. Ye, J. Zhang, S. Jiang, F. Li and Y. Wang, *J. Phys. Chem. B*, 2007, **111**, 5082; (b) S. Hecht and J. M. J. Fréchet, *Angew. Chem., Int. Ed.*, 2001, **40**, 74; (c) B. T. Nguyen, J. E. Gautrot, C. Y. Ji, P. L. Brunner, M. T. Nguyen and X. X. Zhu, *Langmuir*, 2006, **22**, 4799.

- 18 (a) D. Y. Kim, H. N. Cho and C. Y. Kim, *Prog. Polym. Sci.*, 2000, **25**, 1089; (b) C. T. Chen, *Chem. Mater.*, 2004, **16**, 4389.
- 19 (a) K. A. Zachariasse, *Trends Photochem. Photobiol.*, 1994, **3**, 211; (b) J. B. Birks, D. J. Dyson and I. H. Munro, *Proc. R. Soc. London, Ser. A*, 1963, **275**, 575.
- 20 (a) J. A. Mikroyannidis, L. Fenenko and C. Adachi, *J. Phys. Chem. B*, 2006, **110**, 20317; (b) Z.-Q. Liang, Y.-X. Li, J.-X. Yang, Y. Ren and X.-T. Tao, *Tetrahedron Lett.*, 2011, **52**, 1329; (c) C. Tang, F. Liu, Y.-J. Xia, L.-H. Xie, A. Wei, S.-B. Li, Q.-L. Fan and W. Huang, *J. Mater. Chem.*, 2006, **16**, 4074.
- 21 W. Z. Yuan, P. Lu, S. M. Chen, J. W. Y. Lam, Z. M. Wang, Y. Liu, H. S. Kwok, Y. G. Ma and B. Z. Tang, *Adv. Mater.*, 2010, **22**, 2159.
- 22 (a) K. C. Wu, P. J. Ku, C. S. Lin, H. T. Shih, F. I. Wu, M. J. Huang, J. J. Lin, I.-C. Chen and C. H. Cheng, *Adv. Funct. Mater.*, 2008, **18**, 67; (b) T. M. Figueira-Duarte, P. G. Del Rosso, R. Trattnig, S. Sax, E. J. W. List and K. Müllen, *Adv. Mater.*, 2009, **21**, 1.
- 23 (a) F. M. Winnik, *Chem. Rev.*, 1993, **93**, 587; (b) G. Venkataramana and S. Sankararaman, *Org. Lett.*, 2006, **8**, 2739; (c) V. de Halleux, J.-P. Calbert, P. Brocorens, J. Cornil, J.-P. Declercq, J. L. Bredas and Y. Geerts, *Adv. Funct. Mater.*, 2004, **14**, 649; (d) D. Rausch and C. Lambert, *Org. Lett.*, 2006, **8**, 5037; (e) J.-Y. Hu, X. Feng, A. Paudel, H. Tomiyasu, U. Rayhan, P. Thuery, M. R. J. Elsegood, C. Redshaw and T. Yamato, *Eur. J. Org. Chem.*, 2013, 5829; (f) X. Feng, J.-Y. Hu, L. Yi, N. Seto, Z. Tao, C. Redshaw, M. R. J. Elsegood and T. Yamato, *Chem. – Asian J.*, 2012, **7**, 2854; (g) X. Feng, J.-Y. Hu, F. Iwanaga, N. Seto, C. Redshaw and M. R. J. Elsegood, *Org. Lett.*, 2013, **15**, 1318; (h) Y. Qiao, J. Zhang, W. Xua and D. Zhu, *Tetrahedron*, 2011, **67**, 3395.
- 24 I. B. Berlman, *J. Phys. Chem.*, 1970, **74**, 3085.
- 25 (a) J. Luo, Z. Xie, J. W. Y. Lam, L. Cheng, H. Chen, C. Qiu, H. S. Kwok, X. Zhan, Y. Liu, D. Zhu and B. Z. Tang, *Chem. Commun.*, 2001, 1740; (b) Y. Hong, J. W. Y. Lam and B. Z. Tang, *Chem. Commun.*, 2009, 4332; (c) Y. Hong, J. W. Y. Lam and B. Z. Tang, *Chem. Soc. Rev.*, 2011, **40**, 5361; (d) J. Mei, Y. Hong, J. W. Y. Lam, A. Qin, Y. Tang and B. Z. Tang, *Adv. Mater.*, 2014, **26**, 5429; (e) R. Hu, N. L. C. Leung and B. Z. Tang, *Chem. Soc. Rev.*, 2014, **43**, 4494.
- 26 (a) J. Chen, C. C. W. Law, J. W. Y. Lam, Y. Dong, S. M. F. Lo, I. D. Williams, D. Zhu and B. Z. Tang, *Chem. Mater.*, 2003, **15**, 1535; (b) J. Mei, N. L. C. Leung, R. T. K. Kwok, W. Y. Lam and B. Z. Tang, *Chem. Rev.*, 2015, **115**, 11718.
- 27 H. Shi, R. T. K. Kwok, J. Liu, B. Xing, B. Z. Tang and B. Liu, *J. Am. Chem. Soc.*, 2012, **134**, 17972.
- 28 Y. Wang, T. Liu, L. Bu, J. Li, C. Yang, X. Li, Y. Tao and W. Yang, *J. Phys. Chem. C*, 2012, **116**, 15576.
- 29 B. K. An, S. K. Kwon, S. D. Jung and S. Y. Park, *J. Am. Chem. Soc.*, 2002, **124**, 14410.
- 30 Y. Li, F. Li, H. Zhang, Z. Xie, W. Xie, H. Xu, B. Li, F. Shen, L. Ye, M. Hanif, D. Ma and Y. Ma, *Chem. Commun.*, 2007, 231.
- 31 Y. Dong, J. W. Y. Lam, A. Qin, Z. Li, J. Sun, H. H. Y. Sung, I. D. Williams and B. Z. Tang, *Chem. Commun.*, 2007, 40.
- 32 R. Deans, J. Kim, M. R. Machacek and T. M. Swager, *J. Am. Chem. Soc.*, 2000, **122**, 8565.
- 33 K. Perumal, J. A. Garg, O. Blacque, R. Saiganesh, S. Kabilan, K. K. Balasubramanian and K. Venkatesan, *Chem. – Asian J.*, 2012, **7**, 2670.
- 34 J. Wang, J. Mei, R. Hu, J. Z. Sun, A. Qin and B. Z. Tang, *J. Am. Chem. Soc.*, 2012, **134**, 9956.
- 35 (a) N. B. Shustova, B. D. McCarthy and M. Dinca, *J. Am. Chem. Soc.*, 2011, **133**, 20126; (b) Q. Li and Z. Li, *Adv. Sci.*, 2017, **4**, 1600484.
- 36 B. Manimaran, P. Thanasekaran, T. Rajendran, R. J. Lin, I. J. Chang, G. H. Lee, S. M. Peng, S. Rajagopal and K. L. Lu, *Inorg. Chem.*, 2002, **41**, 5323.
- 37 B. Pramanik and D. Das, *J. Phys. Chem. C*, 2018, **122**, 3655.
- 38 (a) R. J. Argauer and C. E. White, *Anal. Chem.*, 1964, **36**, 368; (b) C. A. Royer, *Biophys. J.*, 1995, 1191; (c) M. R. Eftink, *Biophys. J.*, 1994, **66**, 482; (d) S.-J. Kim, F. N. Chowdhury, W. Stryjewski, E. S. Younathan, P. S. Russo and M. D. Barkley, *Biophys. J.*, 1993, **65**, 215.
- 39 Q. Wang, H. Liu, H. Lu, Z. Xu, G. Lai, Z. Li, J. Mack and Z. Shen, *Dyes Pigm.*, 2013, **99**, 771.
- 40 Z. J. Zhao, S. M. Chen, J. W. Y. Lam, P. Lu, Y. C. Zhong, K. S. Wong, H. S. Kwok and B. Z. Tang, *Chem. Commun.*, 2010, **46**, 2221.
- 41 (a) J. Yang, J. Huang, N. Sun, Q. Peng, Q. Li, D. Ma and Z. Li, *Chem. – Eur. J.*, 2015, **21**, 6862; (b) J. Yang, L. Li, Y. Yu, Z. Ren, Q. Peng, S. Ye, Q. Li and Z. Li, *Mater. Chem. Front.*, 2017, **1**, 91.
- 42 X. Feng, C. Qi, H.-T. Feng, Z. Zhao, H. H. Y. Sung, I. D. Williams, R. T. K. Kwok, J. W. Y. Lam, A. Qin and B. Z. Tang, *Chem. Sci.*, 2018, **9**, 5679.
- 43 (a) C. Y. K. Chan, J. W. Y. Lam, J. Zhao, S. M. Chen, P. Lu, H. H. Y. Sung, H. S. Kwok, Y. Ma, I. D. Williams and B. Z. Tang, *J. Mater. Chem. C*, 2014, **2**, 4320; (b) T. Jadhav, J. M. Choi, J. Y. Lee, B. Dhokale and R. Misra, *Org. Electron.*, 2016, **37**, 448; (c) T. Jadhav, B. Dhokale, Y. Patil, S. M. Mobin and R. Misra, *J. Phys. Chem. C*, 2016, **120**, 24030.
- 44 (a) Y. Liu, Y. Tang, N. N. Barashkov, I. S. Irgibaeva, J. W. Y. Lam, R. Hu, D. Birimzhanova, Y. Yu and B. Z. Tang, *J. Am. Chem. Soc.*, 2010, **132**, 13951; (b) L. Chen, Y. Jiang, H. Nie, R. Hu, H. S. Kwok, F. Huang, A. Qin, Z. Zhao and B. Z. Tang, *ACS Appl. Mater. Interfaces*, 2014, **6**, 17215.
- 45 (a) Y. Liu, S. Chen, J. W. Y. Lam, P. Lu, T. K. Kwok, F. Mahtab, H. S. Kwok and B. Z. Tang, *Chem. Mater.*, 2011, **23**, 2536; (b) Z. Zhao, P. Lu, J. W. Y. Lam, Z. Wang, C. Y. K. Chan, H. H. Y. Sung, I. D. Williams, Y. Ma and B. Z. Tang, *Chem. Sci.*, 2011, **2**, 672.
- 46 (a) Z. Zhao, S. Chen, J. W. Y. Lam, Z. Wang, P. Lu, F. Mahtab, H. H. Y. Sung, I. D. Williams, Y. Ma, H. S. Kwok and B. Z. Tang, *J. Mater. Chem.*, 2011, **21**, 7210; (b) W. Wang, T. Lin, M. Wang, T.-X. Liu, L. Ren, D. Chen and S. Huang, *J. Phys. Chem. B*, 2010, **114**, 5983; (c) D. Jana, S. Boxi and B. K. Ghorai, *Dyes Pigm.*, 2013, **99**, 740.
- 47 Z. Zhang, Y. Zhao, R. Zhang, L. Zhang, W. Cheng and Z. H. Ni, *Dyes Pigm.*, 2015, **118**, 95.

- 48 J. Yang, Q. Guo, X. Wen, X. Gao, Q. Peng, Q. Li, D. Ma and Z. Li, *J. Mater. Chem. C*, 2016, **4**, 8506.
- 49 J. Yang, J. Qin, Z. Ren, Q. Peng, G. Xie and Z. Li, *Molecules*, 2017, **22**, 2144.
- 50 S. Son, A. Dodabalapur, A. J. Lovinger and M. E. Galvin, *Science*, 1995, **269**, 376.
- 51 (a) B. Wang, Y. Wang, J. Hua, Y. Jiang, J. Huang, S. Qian and H. Tian, *Chem. – Eur. J.*, 2011, **17**, 2647; (b) B. Gao, H. Wang, Y. Hao, L. Fu, H. Fang, Y. Jiang, L. Wang, Q. Chen, H. Xia, L. Pan, Y. Ma and H. Sun, *J. Phys. Chem. B*, 2010, **114**, 128; (c) Z. Wang, H. Shao, J. Ye, L. Tang and P. Lu, *J. Phys. Chem. B*, 2005, **109**, 19627; (d) J. N. Wilson, M. D. Smith, V. Enkelmann and U. H. F. Bunz, *Chem. Commun.*, 2004, 1700; (e) C. J. Bhongale, C.-S. Chang, C.-S. Lee, E. W.-G. Diao and C.-S. Hsu, *J. Phys. Chem. B*, 2005, **109**, 13472; (f) H. Tong, Y. Dong, M. Häußler, Y. Hong, J. W. Y. Lam, H. H.-Y. Sung, I. D. Williams, H. S. Kwok and B. Z. Tang, *Chem. Phys. Lett.*, 2006, **428**, 326; (g) P. An, Z. Shi, W. Dou, X. Cao and H. Zhang, *Org. Lett.*, 2010, **12**, 4364.
- 52 Y.-X. Li, G.-X. Sun, J.-L. Miao, Y. Nie, Z.-W. Zhang and X.-T. Tao, *Tetrahedron Lett.*, 2013, **54**, 3263.
- 53 Y.-X. Li, X.-F. Yang, J.-L. Miao and G.-X. Sun, *J. Phys. Chem. C*, 2016, **120**, 21722.
- 54 J. Katla, H. R. Bhat, P. C. Jha, P. S. Ghalsasi and S. Kanvah, *ChemistrySelect*, 2017, **2**, 1902.
- 55 A. D. Aleo, A. Karapetyan, V. Heresanu, M. Giorgi and F. Fages, *Tetrahedron*, 2015, **71**, 2255.
- 56 A. Karuppusamy, T. Vandana and P. Kannan, *J. Photochem. Photobiol., A*, 2017, **345**, 11.
- 57 (a) D. A. Pearce, N. Jotterand, I. S. Carrico and B. Imperiali, *J. Am. Chem. Soc.*, 2001, **123**, 5160; (b) M. Mameli, M. C. Aragoni, M. Arca, C. Caltagirone, F. Demartin, G. Farruggia, G. D. Filippo, F. A. Devillanova, A. Garau, F. Isaia, V. Lippolis, S. Murgia, L. Prodi, A. Pintus and N. Zaccheroni, *Chem. – Eur. J.*, 2010, **16**, 919; (c) Y. Bao, B. Liu, H. Wang, F. Du and R. Bai, *Anal. Methods*, 2011, **3**, 1274; (d) Y. Bao, B. Liu, F. Du, J. Tian, H. Wang and R. Bai, *J. Mater. Chem.*, 2012, **22**, 5291.
- 58 M. Shellaiah, Y.-H. Wu, A. Singh, M. V. R. Raju and H.-C. Lin, *J. Mater. Chem. A*, 2013, **1**, 1310.
- 59 (a) D. T. McQuade, A. E. Pullen and T. M. Swager, *Chem. Rev.*, 2000, **100**, 2537; (b) G. Aragay, J. Pons and A. Merkoci, *Chem. Rev.*, 2011, **111**, 3433.
- 60 A. Kathiravan, K. Sundaravel, M. Jaccob, G. Dhinakaran, A. Rameshkumar, D. A. Ananth and T. Sivasudha, *J. Phys. Chem. B*, 2014, **118**, 13573.
- 61 Y. Li, W. Huang, J. Yong, S. Huang, Y. Li, Y. Liu and D. Wu, *New J. Chem.*, 2018, **42**, 12644.
- 62 Q. Chai, J. Wei, B. Bai, H. Wang and M. Li, *Dyes Pigm.*, 2018, **152**, 93.
- 63 M. Shellaiah, T. Simon, V. Srinivasadesikan, C.-M. Lin, K. W. Sun, F.-H. Ko, M.-C. Lin and H.-C. Lin, *J. Mater. Chem. C*, 2016, **4**, 2056.
- 64 J. E. Kwon and S. Y. Park, *Adv. Mater.*, 2011, **23**, 3615.
- 65 S. Samanta, U. Manna and G. L. Das, *New J. Chem.*, 2017, **41**, 1064.
- 66 M. Shyamal, P. Mazumdar, S. Maity, G. P. Sahoo, G. Salgado-Morán and A. Misra, *J. Phys. Chem. A*, 2016, **120**, 210.
- 67 M. Shyamal, S. Maity, P. Mazumdar, G. P. Sahoo, R. Maity and A. Misra, *J. Photochem. Photobiol., A*, 2017, **342**, 1.
- 68 T. Jadhav, B. Dhokale, S. M. Mobin and R. Misra, *J. Mater. Chem. C*, 2015, **3**, 9981.
- 69 Y. Liu, T. Shan, L. Yao, Q. Bai, Y. Guo, J. Li, X. Han, W. Li, Z. Wang, B. Yang, P. Lu and Y. Ma, *Org. Lett.*, 2015, **17**, 6138.
- 70 Y. Sagara, T. Mutai, I. Yoshikawa and K. Araki, *J. Am. Chem. Soc.*, 2007, **129**, 1520.
- 71 (a) Y. Sagara and T. Kato, *Angew. Chem., Int. Ed.*, 2008, **47**, 5175; (b) M. J. Teng, X. R. Jia, S. Yang, X. F. Chen and Y. Wei, *Adv. Mater.*, 2012, **24**, 1255; (c) M. J. Teng, X. R. Jia, X. F. Chen and Y. Wei, *Angew. Chem., Int. Ed.*, 2012, **51**, 6398; (d) T. S. Wang, N. Zhang, K. Zhang, J. W. Dai, W. Bai and R. K. Bai, *Chem. Commun.*, 2016, **52**, 9679.
- 72 L. Ji, A. Lorbach, R. M. Edkins and T. B. Marder, *J. Org. Chem.*, 2015, **80**, 5658.
- 73 (a) L. Zöphel, D. Beckmann, V. Enkelmann, D. Chercka, R. Rieger and K. Müllen, *Chem. Commun.*, 2011, **47**, 6960; (b) T. Yamato, M. Fuimoto, A. Miyazawa and K. Matsuo, *J. Chem. Soc., Perkin Trans. 1*, 1997, 1201; (c) B. Kohl, F. Rominger and M. Mastalerz, *Angew. Chem., Int. Ed.*, 2015, **54**, 6051; (d) H. M. Duong, M. Bendikov, D. Steiger, Q. Zhang, G. Sonmez, J. Yamada and F. Wudl, *Org. Lett.*, 2003, **5**, 4433; (e) K. Baumgärtner, K. A. L. M. Chinchu, A. Dreuw, F. Rominger and M. Mastalerz, *Angew. Chem., Int. Ed.*, 2016, **55**, 1.
- 74 (a) L. Zöphel, V. Enkelmann and K. Müllen, *Org. Lett.*, 2013, **15**, 804; (b) S. N. Keller, N. L. Veltri and T. C. Sutherland, *Org. Lett.*, 2013, **15**, 4798; (c) M. Piotrowicz, J. Zakrzewski, R. Métivier, A. Brosseau, A. Makal and K. Woźniak, *J. Org. Chem.*, 2015, **80**, 2573.
- 75 S. Sasaki, S. Suzuki, K. Igawa, K. Morokuma and G.-I. Konishi, *J. Org. Chem.*, 2017, **82**, 6865.
- 76 V. Kachwal, P. Alam, H. R. Yadav, S. S. Pasha, A. R. Choudhury and I. R. Laskar, *New J. Chem.*, 2018, **42**, 1133.
- 77 U. K. Das and P. Dastidar, *J. Chem. Sci.*, 2014, **126**, 1357.
- 78 K. Nishino, H. Yamamoto, K. Tanaka and Y. Chujo, *Org. Lett.*, 2016, **18**, 4064.
- 79 V. Vij, V. Bhalla and M. Kumar, *Chem. Rev.*, 2016, **116**, 9565.
- 80 (a) C.-Z. Wang, Y. Noda, C. Wu, X. Feng, P. Venkatesan, H. Cong, M. R. J. Elsegood, T. G. Warwick, S. J. Teat, C. Redshaw and T. Yamato, *Asian J. Org. Chem.*, 2018, **7**, 444; (b) C. Wang and Z. Li, *Mater. Chem. Front.*, 2017, **1**, 2174.
- 81 (a) Y. Niko, S. Kawauchi, S. Otsu, K. Tokumaru and G.-I. Konishi, *J. Org. Chem.*, 2013, **78**, 3196; (b) Y. Niko, Y. Cho, S. Kawauchi and G.-I. Konishi, *RSC Adv.*, 2014, **4**, 36480.
- 82 (a) A. Matsumoto, M. Suzuki, D. Kuzuhara, J. Yuasa, T. Kawai, N. Aratani and H. Yamada, *Chem. Commun.*, 2014, **50**, 10956; (b) H. Lee, B. Kim, S. Kim, J. Kim, J. Lee, H. Shin, J. Lee and J. Park, *J. Mater. Chem. C*, 2014, **2**, 4737.
- 83 (a) J. N. Moorthy, P. Natarajin, P. Venkatakrishnan, D. F. Huang and T. J. Chow, *Org. Lett.*, 2007, **9**, 5215; (b) M. Y. Lo, C. G. Zhen, M. Lauters, G. E. Jabbour and

- A. Sellinger, *J. Am. Chem. Soc.*, 2007, **129**, 5808; (c) R. D. Xia, W. Y. Lai, P. A. Levermore, W. Huang and D. D. C. Bradley, *Adv. Funct. Mater.*, 2009, **19**, 2844; (d) Z. J. Zhao, J. H. Li, X. P. Chen, X. M. Wang, P. Lu and Y. Yang, *J. Org. Chem.*, 2009, **74**, 383.
- 84 (a) H. J. Zhang, Y. Wang, K. Z. Shao, Y. Q. Liu, S. Y. Chen, W. F. Qiu, X. B. Sun, T. Qi, Y. Q. Ma, G. Yu, Z. M. Su and D. B. Zhu, *Chem. Commun.*, 2006, 755; (b) M. J. Sienkowska, J. M. Farrar, F. Zhang, S. Kusuma, P. A. Heiney and P. Kaszynski, *J. Mater. Chem.*, 2007, **17**, 1399; (c) F. Liu, W. Y. Lai, C. Tang, H. B. Wu, Q. Q. Chen, B. Peng, W. Wei, W. Huang and Y. Cao, *Macromol. Rapid Commun.*, 2008, **29**, 659.
- 85 X. Feng, Z. Xu, Z. Hu, C. Qi, D. Luo, X. Zhao, Z. Mu, C. Redshaw, J. W. Y. Lam, D. Ma and B. Z. Tang, *J. Mater. Chem. C*, 2019, **7**, 2283.
- 86 N. J. Jeon, J. Lee, J. H. Noh, M. K. Nazeeruddin, M. Grätzel and S. Seok, *J. Am. Chem. Soc.*, 2013, **135**, 19087.
- 87 (a) J. R. Lakowicz, *Principles of Fluorescence Spectroscopy*, Springer, New York, 2006; (b) W. Z. Yuan, S. Chen, J. W. Y. Lam, C. Deng, P. Lu, H. H. Y. Sung, I. D. Williams, H. S. Kwok, Y. Zhang and B. Z. Tang, *Chem. Commun.*, 2011, **47**, 11216; (c) C. Shi, Z. Guo, Y. Yan, S. Zhu, Y. Xie, Y. S. Zhao, W. Zhu and H. Tian, *ACS Appl. Mater. Interfaces*, 2013, **5**, 192.
- 88 (a) X. Gao and Z. Zhao, *Sci. China: Chem.*, 2015, **58**, 947; (b) K. Oniwa, H. Kikuchi, H. Shimotani, S. Ikeda, N. Asao, Y. Yamamoto, K. Tanigaki and T. Jin, *Chem. Commun.*, 2016, **52**, 4800; (c) F. Moggia, C. Videlot-Ackermann, J. Ackermann, P. Raynal, H. Brisset and F. Fages, *J. Mater. Chem.*, 2006, **16**, 2380; (d) J. E. Anthony, *Angew. Chem., Int. Ed.*, 2008, **47**, 452.
- 89 J. Zaumseil and H. Sirringhaus, *Chem. Rev.*, 2007, **107**, 1296.
- 90 H. Ju, K. Wang, J. Zhang, H. Geng, Z. Liu, G. Zhang, Y. Zhao and D. Zhang, *Chem. Mater.*, 2017, **29**, 3580.
- 91 (a) X. J. Tang, L. Z. Wu, L. P. Zhang and C. H. Tung, *Chem. Phys. Lett.*, 2002, **356**, 573; (b) G. S. He, T. C. Lin, V. K. S. Hsiao, A. N. Cartwright, P. N. Prasad, L. V. Natarajan, V. P. Tondiglia, R. Jakubiak, R. A. Vaia and T. J. Bunning, *Appl. Phys. Lett.*, 2003, **83**, 2733.
- 92 J. Huang, Y. Wu, Y. Chen, Z. Zhu, X. Yang, C. J. Yang, K. Wang and W. Tan, *Angew. Chem., Int. Ed.*, 2011, **50**, 401.
- 93 J. H. Clint, *Surfactant Aggregation: Blacki*, Chapman and Hall, New York, 1991.
- 94 (a) S. K. Hait, P. R. Majhi, A. Blume and S. P. Moulik, *J. Phys. Chem. B*, 2003, **107**, 3650; (b) M. Prasad, A. MacDonald, R. Palepu and S. P. Moulik, *J. Surf. Sci. Technol.*, 2003, **19**, 125.
- 95 (a) G. B. Ray, I. Chakraborty and S. P. Moulik, *J. Colloid Interface Sci.*, 2006, **294**, 248; (b) K. Kalyansundaran and J. K. Thomas, *J. Am. Chem. Soc.*, 1977, **99**, 2039; (c) J. Aguiar, P. Carpana, J. A. Molina-Boliver and C. C. Ruiz, *J. Colloid Interface Sci.*, 2003, **258**, 116.
- 96 (a) Y. Sagara, S. Yamane, M. Mitani, C. Weder and T. Kato, *Adv. Mater.*, 2016, **28**, 1073; (b) Y. Q. Dong, J. W. Y. Lam, B. Z. Tang and J. Phys. Chem. Lett., 2015, **6**, 3429; (c) J. Picard, *Science*, 1939, **89**, 460; (d) N. C. Eddingsaas and K. S. Suslick, *Nature*, 2006, **444**, 163.
- 97 Y.-B. Gong, P. Zhang, Y.-R. Gu, J.-Q. Wang, M.-M. Han, C. Chen, X.-J. Zhan, Z.-L. Xie, B. Zou, Q. Peng, Z.-G. Chi and Z. Li, *Adv. Opt. Mater.*, 2018, **6**, 1800198.
- 98 J. Wang, X. Gu, P. Zhang, X. Huang, X. Zheng, M. Chen, H. Feng, R. T. K. Kwok, J. W. Y. Lam and B. Z. Tang, *J. Am. Chem. Soc.*, 2017, **139**, 16974.
- 99 (a) M. P. Fernández-Liencres, T. Peña-Ruiz, J. M. Granadino-Roldan, M. Moral, A. Valenzuela-Pereira, A. Garzón-Ruiz and A. Navarro, *J. Phys. Chem. C*, 2018, **122**, 12002; (b) A. G. Crawford, A. D. Dwyer, Z. Liu, A. Steffen, A. Beeby, L.-O. Pålsson, D. J. Tozer and T. B. Marder, *J. Am. Chem. Soc.*, 2011, **133**, 13349.
- 100 C. Adachi, M. A. Baldo, M. E. Thompson and S. R. Forrest, *J. Appl. Phys.*, 2001, **90**, 5048.
- 101 K. Suzuki, S. Kubo, K. Shizu, T. Fukushima, A. Wakamiya, Y. Murata, C. Adachi and H. Kaji, *Angew. Chem., Int. Ed.*, 2015, **127**, 15446.
- 102 (a) W. Zhao, Z. He, Q. Peng, J. W. Y. Lam, H. Ma, Z. Qiu, Y. Chen, Z. Zhao, Z. Shuai, Y. Dong and B. Z. Tang, *Nat. Commun.*, 2018, **9**, 3044; (b) Z. Yang, Z. Mao, X. Zhang, D. Ou, Y. Mu, Y. Zhang, C. Zhao, S. Liu, Z. Chi, J. Xu, Y.-C. Wu, P.-Y. Lu, A. Lien and M. R. Bryce, *Angew. Chem., Int. Ed.*, 2016, **55**, 2181.
- 103 J. Huang, H. Nie, J. Zeng, Z. Zhuang, S. Gan, Y. Cai, J. Guo, S.-J. Su, Z. Zhao and B. Z. Tang, *Angew. Chem., Int. Ed.*, 2017, **56**, 12971.
- 104 (a) P. Rajamalli, N. Senthilkumar, P. Gandeepan, P.-Y. Huang, M.-J. Huang, C.-Z. Ren-Wu, C.-Y. Yang, M.-J. Chiu, L.-K. Chu, H.-W. Lin and C.-H. Cheng, *J. Am. Chem. Soc.*, 2016, **138**, 628; (b) G. Xie, J. Luo, M. Huang, T. Chen, K. Wu, S. Gong and C. Yang, *Adv. Mater.*, 2017, **29**, 1604223.
- 105 L. L. Liu, J. Zhou, R. Andrews and D. W. Stephan, *J. Am. Chem. Soc.*, 2018, **14**, 7466.