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Controllable synthesis of a Na-enriched $\text{Na}_4\text{V}_2(\text{PO}_4)_3$ cathode for high-energy sodium-ion batteries: a redox-potential-matched chemical sodiation approach†

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Exploring a sodium-enriched cathode (*i.e.* $\text{Na}_4\text{V}_2(\text{PO}_4)_3$, which differs from its traditional stoichiometric counterpart $\text{Na}_3\text{V}_2(\text{PO}_4)_3$ that can provide extra endogenous sodium reserves to mitigate the irreversible capacity loss of the anode material (*i.e.* hard carbon), is an intriguing presodiation method for the development of high energy sodium-ion batteries. To meet this challenge, herein, we first propose a redox-potential-matched chemical sodiation approach, utilizing phenazine-sodium (PNZ-Na) as the optimal reagent to sodiate the $\text{Na}_3\text{V}_2(\text{PO}_4)_3$ precursor into Na-enriched $\text{Na}_4\text{V}_2(\text{PO}_4)_3$. The spontaneous sodiation reaction enables a fast reduction of one-half V ions from V^{3+} to V^{2+} , followed by the insertion of one Na^+ ion into the NASICON framework, which only takes 90 s to obtain the phase-pure $\text{Na}_4\text{V}_2(\text{PO}_4)_3$ product. When paired with a hard carbon anode, the resulting $\text{Na}_4\text{VP}||\text{HC}$ full cell exhibits a high energy density of 251 W h kg^{-1} , which is 58% higher than that of 159 W h kg^{-1} for the $\text{Na}_3\text{VP}||\text{HC}$ control cell. Our chemical sodiation methodology provides an innovative approach for designing sodium-rich cathode materials and could serve as an impetus to the development of advanced sodium-ion batteries.

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Introduction

Lithium-ion batteries (LIBs) are currently the dominant energy storage technology for portable and automotive applications; however, the high cost and geographical constraints of lithium resources make it impossible to sustain and meet the exploding market demands.¹ As an alternative to LIBs, sodium-ion batteries (SIBs) are gaining increased recognition as next generation low cost energy storage devices, owing to the natural abundance of sodium resources ($\sim 2.5\%$ in the Earth's crust) and their similar electrochemical properties to LIBs.^{2,3} A series of high-performance sodium storage electrode materials have been investigated to pursue the ever-growing demand for higher energy density.⁴ Cathode materials such as layered oxides,⁵ polyanionic compounds,^{6,7} and Prussian blue analogues,⁸ as well as anode materials such as hard carbons,⁹ conversion

materials,¹⁰ and alloy compounds,^{11,12} have all demonstrated high specific capacity and excellent cycle stability.

However, the low initial coulombic efficiency (ICE), an intrinsic problem of anode materials, is generally ignored by researchers but seriously impedes the improvement of the energy density of SIBs.^{13,14} The low ICE signifies that a large number of active sodium ions are irreversibly consumed to form a solid electrolyte interphase (SEI) layer on the anode during the first cycle.¹⁵ This adverse effect is not as strongly felt in half cells, since the Na metal counter electrode can provide ample Na sources. However, as for Na-ion full cells, where the Na-containing cathode is the only reserve of active sodium ions, such a considerable Na^+ loss dramatically reduces the overall energy density and shortens the lifespan of full cells. For instance, hard carbon (HC) anodes usually have a low ICE of only 60–80%, which leads to nearly one-third of capacity loss during the initial charge process.¹⁶

Numerous research efforts have been committed to resolving this issue by offering an additional Na source to counteract the initial capacity loss (ICL) of anodes, known as presodiation.^{17,18} In the early stage, presodiation has been carried out by the operation of highly active Na metal. Examples include direct contact of the anode electrodes with Na metal disks¹⁹ or assembling them into a primary cell.²⁰ In these cases, the irreversible capacity can be readily dismissed *via* the self-discharge

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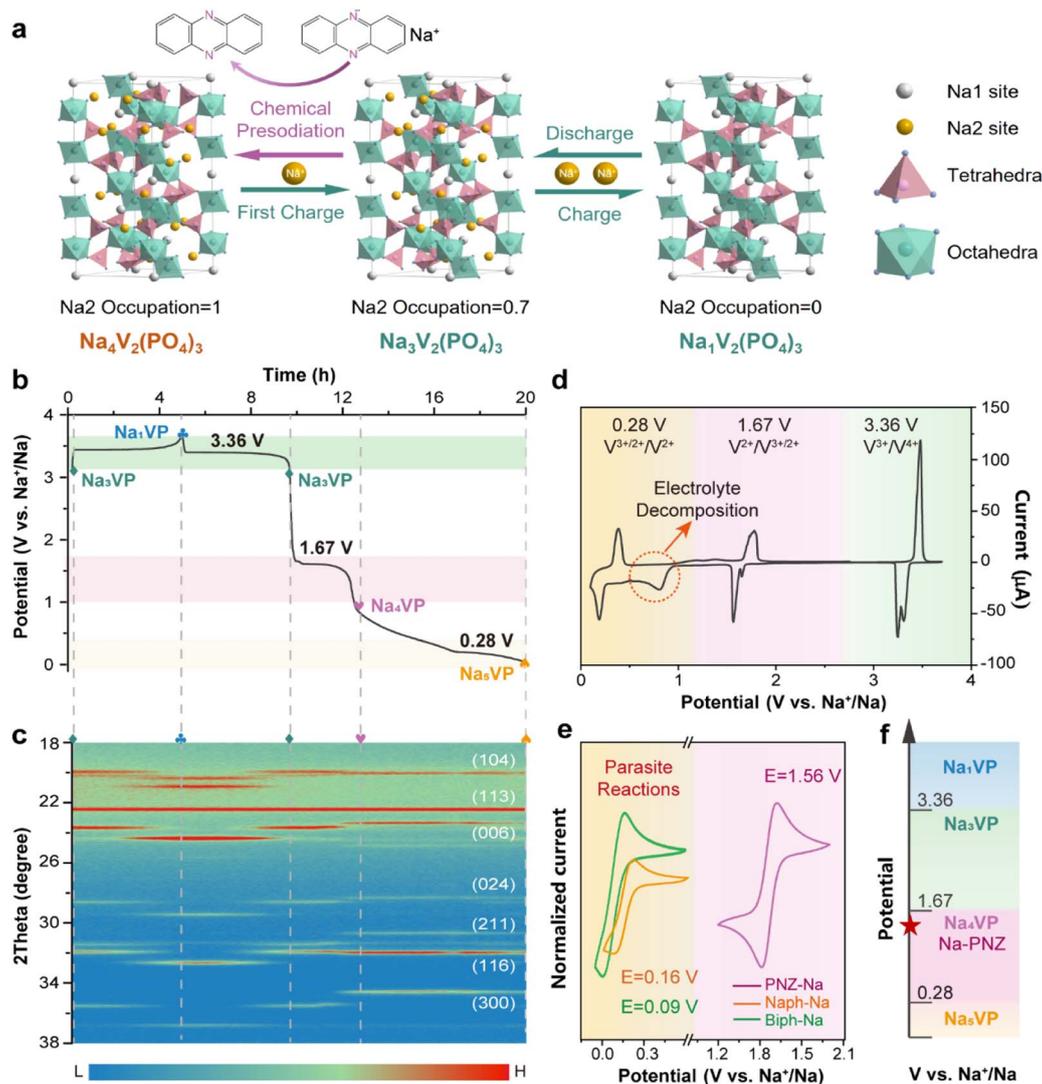


Fig. 1 (a) Schematic illustration of chemical presodiation and the subsequent cycling process of the Na₄VP cathode. (b) Initial charge–discharge profile of Na₃VP and (c) the corresponding *in situ* XRD counter map. (d) Cyclic voltammogram profiles of the Na₃VP cathode in the potential range of 0.1–3.7 V vs. Na⁺/Na. (e) Cyclic voltammogram profiles of different sodiation reagents: phenazine-sodium (PNZ-Na), naphthalene-sodium (Naph-Na), and biphenyl-sodium (Biph-Na) at a scan rate of 50 mV s⁻¹. (f) Illustration diagram of the potential matching principle for chemical sodiation of Na₃VP.

and 34.7° with gradually increasing intensity, which is in good agreement with previous reports.^{27,41} This demonstrates that the two phases coexist and correspond directly to the voltage plateau. Upon further sodiation, a new plateau appears at about 0.28 V (vs. Na⁺/Na) while diffraction peaks at 23.5° and 28.5° fade away gradually, accompanying the formation of Na₅VP.²⁸ Operando XRD confirms the feasibility of the presodiation strategy to produce sodium-enriched phases from Na₃VP and reveal the corresponding structure information.

The cyclic voltammogram (CV) profile (Fig. 1d) obtained between 0.1 and 3.7 V (vs. Na⁺/Na) reconfirms the accessibility of Na₄VP and Na₅VP. Three pairs of cathodic/anodic peaks located at ~0.20/0.36, ~1.56/1.78 and ~3.30/3.48 V are in good agreement with plateaus (observed in the galvanostatic charge/discharge profiles (eqn (1)–(3)). Notably, a broad irreversible peak is observed between 0.3 and 0.9 V, implying aggressive

electrolyte decomposition. During the cycles (Fig. S2†), all the peak currents prominently diminish and the polarization gradually increases, owing to the structure deterioration of active materials and the accumulation of the SEI on the electrode surface. In contrast, the CV curves between 1.0 and 4.0 V almost completely overlapped after the initial cycle (Fig. S3†), indicating the highly reversible phase transition among Na₁VP, Na₃VP, and Na₄VP. The split of the cathodic peak in the first cycle may be ascribed to the minor structure rearrangement of the active material.⁴² The narrower CV peaks also imply a faster Na⁺ insertion process for Na₄VP than Na₅VP. It is important to note that Na₅VP exhibits low initial coulombic efficiency (ICE_{cathode} is defined as the ratio of sodiation capacity to desodiation capacity of a cathode), rendering it incompatible with HC anodes (Fig. S4†). As mentioned above, the typical ICE_{anode} (defined as the ratio of desodiation capacity to sodiation



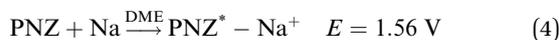
capacity of an anode) of HC anodes is only 60–80%. When Na₅VP with an ICE of merely 40% matches with HC anodes, surplus Na⁺ will be introduced into the full cells and form uncontrollable sodium dendrites on the anode side, resulting in safety concerns. Therefore, Na₄VP has a suitable ICE (~64%) and excellent structural stability, denoting its superiority for active sodium compensation in full cells.

As displayed in Fig. 1e, currently reported aromatic sodium reagents, such as Biph-Na and Naph-Na, demonstrate ultra-low redox potentials of 0.16 V and 0.09 V vs. Na⁺/Na for cathode materials (lower than the Na₅VP formation potential of 0.28 V vs. Na⁺/Na). In order to prevent over-sodiation from yielding Na₅VP or other parasite reactions, the treatment duration with sodium reagents must be strictly controlled, which is not conducive to industrial batch application. Therefore, it is of utmost importance to develop appropriate sodium reagents to achieve a refined precise and controllable preparation of Na₄VP.

Na⁺ insertion/extraction reaction of Na_xVP:



Preparation of the PNZ-Na reagent:



Precise chemical sodiation of Na₃VP enabled by the PNZ-Na reagent:



In accordance with the potential-matched chemical sodiation strategy (Fig. 1f), phenazine sodium (PNZ-Na) has been elaborately chosen as a suitable sodiation reagent due to its appropriate potential. The PNZ-Na reagent displays a pair of reversible redox peaks at 1.51 and 1.62 V (vs. Na⁺/Na, Fig. 1e), corresponding to the PNZ/PNZ⁺ redox couple (eqn (4)). PNZ-Na exhibits the highest potential as a sodiation reagent reported so far,^{26,32,33,43} benefiting from the strong electron-withdrawing effect of nitrogen atoms and the efficient charge dispersion provided by the additional six-membered ring.^{44,45} Hence, the end-point of the reaction is effectively clamped by the PNZ-Na reagent potential, thus preventing over-sodiation. During the chemical presodiation process, reductive PNZ-Na concurrently donates Na⁺ and e⁻ for Na₃VP while also leading to the recovery of intrinsic phenazine molecules (eqn (5)). The sodium ions diffuse into the Na₃VP lattice and locate at the partially occupied Na2 site to form a pure Na₄VP phase.

Controllable synthesis of Na-enriched Na₄VP via the chemical sodiation strategy

The chemical presodiation of the Na₃VP cathode is conducted simply by immersing the electrode into PNZ-Na solution for a particular time, followed by DME rinse to remove the residual

PNZ-Na. Phenazine and sodium metal are mixed in dimethoxyethane (DME) solvent in equal amounts to obtain the PNZ-Na reagent, as described in eqn (4). The strong electron affinity of heterocyclic aromatics causes the outermost electrons of Na metal to transfer spontaneously to the conjugated phenazine rings, producing grape-purple colored PNZ⁻ radicals (Fig. 2a left).^{46,47} Once the Na₃VP electrode is immersed in the PNZ-Na solution, the sodiation reaction proceeds through electron transfer from the PNZ⁻ radical to Na₃VP along with Na⁺ accommodation to form Na₄VP (eqn (5)). Following the reaction, the purple PNZ⁻ radical dissolved in DME is converted to a pristine PNZ molecule, resulting in the PNZ-Na solution fading (Fig. 2a right).

To probe the reaction pathways, we prepared presodiated Na_xVP (3 ≤ x ≤ 4) electrodes with various immersion times of 30–120 s. As the reaction time increases, the emergence of new diffraction peaks at 30.8° and 34.7°, along with growing intensity, is attributable to the Na₄VP phase, while the peaks of Na₃VP at 31.6° (116) and 35.6° (300) weaken (Fig. 2b). This outcome reveals that two phases, Na₃VP and Na₄VP, coexist during the presodiation process (30 s and 60 s), which is consistent with *in situ* XRD and prior reports.^{26,28} Notably, Na₃VP is completely converted to Na₄VP after presodiation for more than 90 s. (Fig. S5†) The XRD pattern of chemically presodiated Na₄VP is identical to that of electrochemically sodiated Na₄VP, implying a rapid Na⁺ migration in the cathode and a uniform sodiation process achieved by chemical presodiation. The inductively coupled plasma atomic emission spectroscopy (ICP-OES) tests accurately determine the Na content in different samples (Fig. 2c). The samples with treatment time of 30, 60, and over 90 s are accordingly denoted as Na_{3.37}VP, Na_{3.68}VP, and Na₄VP.

The presodiation degree of Na_xVP achieved *via* chemical treatment is also evaluated using electrochemical performance in sodium half-coin cells. The initial charge curves of Na_xVP, including ICE and plateau capacity at about 1.67 V that correspond to Na₄VP, are displayed in Fig. 2d and e. The open circuit voltage (OCV vs. Na⁺/Na) dramatically drops during the biphasic process from Na₃VP to Na₄VP, after which it remains stable at around 1.67 V. The presodiated Na_xVP exhibits a continuous increase in initial charge capacity (*i.e.*, Na extraction), whereas the ICE decreases with increasing sodiation time, indicating a sustained Na⁺ insertion reaction (Fig. S6†). Consequently, the chemical presodiation process continues until Na₃VP diminishes, as reflected by a constant ICE of ~64% and an increased specific capacity of ~56 mA h g⁻¹, respectively. Notably, no unnecessary parasitic reactions are introduced due to the identical charge-discharge curves and ICE between chemical and electrochemical sodiation (Fig. S7a†). The redox potentials between the reductive agent and Na₃VP are well-matched, thereby avoiding the issues of over-sodiation, inhomogeneity, and inconsistency during scaling up for mass production. Even with excessive treatment time (30 min) and reagent dosage, there is no structural damage to the cathode materials, demonstrating the success of our chemical presodiation strategy based on the redox-potential matched principle (Fig. S7b†). The sodiation reaction self-terminates once all



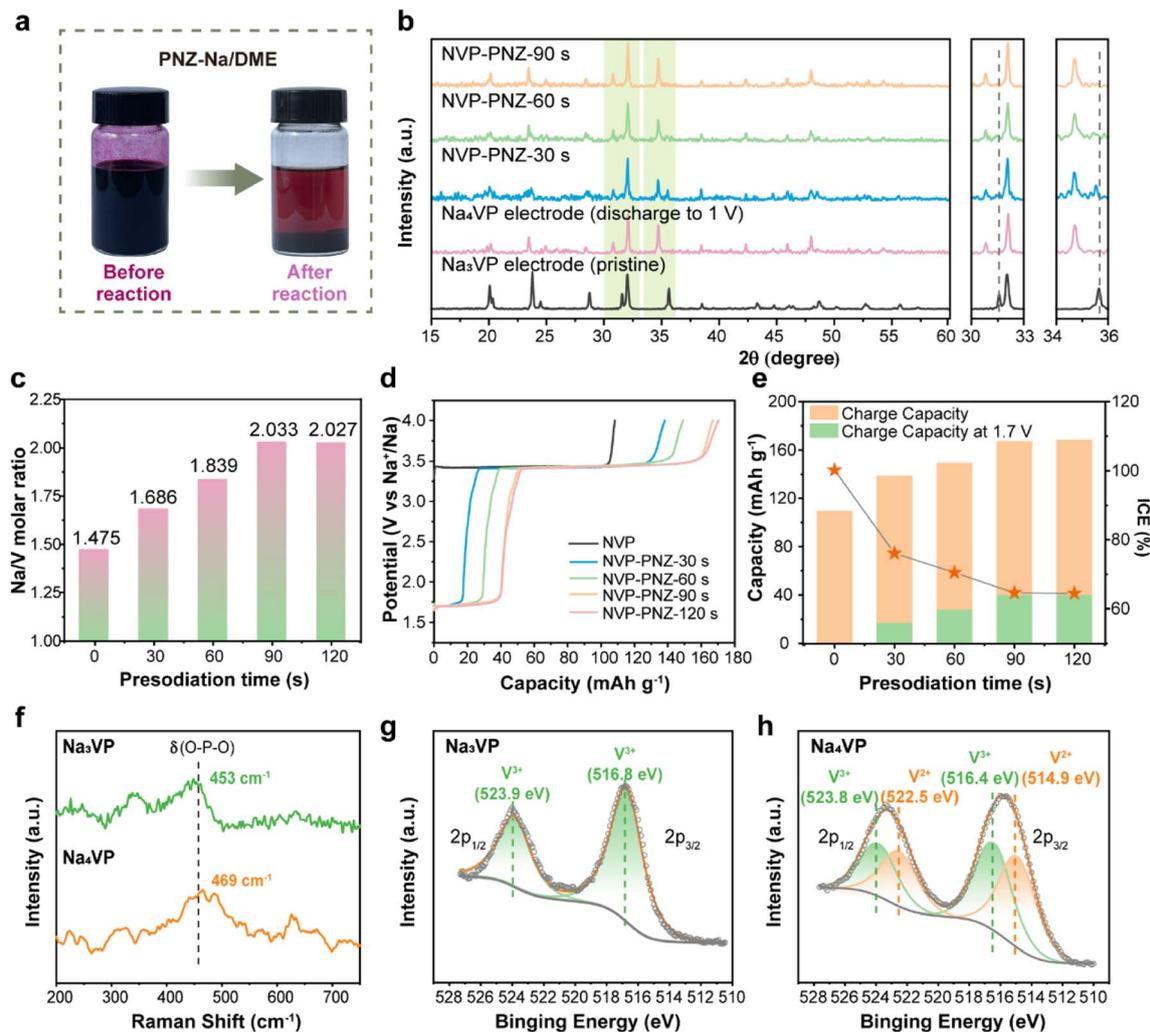


Fig. 2 (a) Optical photographs of the phenazine sodium reagent before and after reaction with Na_3VP . (b) XRD patterns, (c) ICP results, (d) initial charge curve, and (e) initial charge capacity and ICE of the Na_xVP electrodes with different presodiation times. (f) Raman spectroscopy spectra of Na_3VP and Na_4VP . XPS V 2p spectra of (g) Na_3VP and (h) Na_4VP .

Na_3VP transforms into sodium-rich Na_4VP due to the elaborately screened sodiation reagents. We further conducted the *in situ* XRD test of the PNZ-Na presodiated Na_4VP electrode (Fig. S8 and S9[†]). The variation of diffraction peaks is consistent with the electrochemical sodiation process, and there is no impurity phase introduced by presodiation treatment. It is a highly reversible transformation between Na_4VP and Na_3VP and reconfirms that presodiation has no adverse effect on the lattice structure. Therefore, the reliability of the chemical presodiation strategy is firmly authenticated.

Raman spectra and X-ray photoelectron spectroscopy (XPS) are further employed in order to explore the effect of presodiation treatment on the structure of the Na_3VP cathode. The Raman spectra shown in Fig. 2f indicate that the symmetric bending vibration peak of the PO_4 tetrahedron in Na_3VP at 453 cm^{-1} is slightly blue-shifted after presodiation treatment. This result suggests the enhancement of P–O bond energy and boosting ionicity of the Na_4VP lattice.⁴⁸ The valence states of vanadium ions are verified by the XPS results. The survey XPS

spectrum (Fig. S10[†]) confirms the presence of the same chemical composition elements in Na_3VP and Na_4VP . The binding energy of P 2p shifts toward the higher energy direction (from 133.4 eV to 133.7 eV), reconfirming the stronger P–O bonds (Fig. S11[†]). As seen in Fig. 2g and h, the V 2p peaks broaden after presodiation and reveal the coexistence of mixed V^{3+} and V^{2+} ions in the lattice. The two peaks of Na_3VP at 516.8 eV and 523.9 eV coincide with V $2p_{3/2}$ and V $2p_{1/2}$ of V^{3+} , which means only trivalent vanadium is observed in the pristine electrode.⁴⁹ Except for the peaks above, new peaks at 514.9 eV and 522.5 eV can be ascribed to the contribution of V^{2+} in Na_4VP .²⁶ Further quantitative analysis of the V 2p spectrum indicates a $\text{V}^{3+}/\text{V}^{2+}$ ratio of $0.49/0.51$ based on the integral area of deconvoluted peaks, which agrees well with the valence distribution for Na_4VP . All the characterization studies above demonstrate that the PNZ-Na solution can readily and controllably presodiate the Na_3VP cathode in a relatively short time.

The morphological evolution of the Na_3VP and Na_4VP samples is investigated by scanning electron microscopy (SEM),



focused ion beam (FIB), and high-resolution transmission electron microscopy (HRTEM). The results show that both Na₃VP and Na₄VP exhibit a regular spherical morphology, with a diameter ranging from 5–15 μm, consisting of primary particles that are densely packed to form secondary particles (Fig. 3a and d). Apparently, the morphology characteristic of sodiated Na₄VP is not affected by the chemical treatment and helps to retain the electrochemical performance (Fig. S12†). The energy-dispersive X-ray (EDX) analysis reconfirms the Na/V ratios of Na₃VP and Na₄VP to be 1.49 and 1.99, respectively (Fig. S13†). Cross sections are obtained by FIB-SEM to monitor microstructural damage, while no microcracks are observed after presodiation treatment (Fig. 3b and e). It should be noted that the internal voids in the particles are formed by primary particle agglomeration rather than structure degradation. HRTEM images in Fig. 3c reveal well-defined lattice fringes of 0.373 nm ascribed to the (113) planes in the Na₃VP lattice. Likewise, we observe characteristic lattice fringes of 0.378 nm (Fig. 3f), corresponding to the emerging 23.5° diffraction peak in the Na₄VP lattice (Fig. 2b).

²³Na/³¹P solid state nuclear magnetic resonance (ssNMR) offers in-depth insights into local structural transformation while probing the oxidation states of the neighboring vanadium ions. For ³¹P spectra, the signals are caused by the unpaired electronic spin transfer from TM d orbitals to the ³¹P nuclei through chemical bonds.⁵⁰ Fig. S14† presents the room temperature ³¹P spectra for Na₃VP and Na₄VP. In the Na₃VP lattice, a broad asymmetric peak at 2000–3500 ppm reflects a twisted environment for the ³¹P nuclei.⁵¹ While in Na₄VP, the corresponding peak with decreasing intensity reveals the reducing dipolar interaction between unpaired electron spins of V and P nuclei. In contrast, a new broad peak at 3500–5000 ppm is observed in Na₄VP, which is attributed to the

complicated interaction between V²⁺ and adjacent phosphate groups.⁵⁰ This phenomenon is because the V²⁺ (t_{2g}³e_g⁰) owning three unpaired electrons is more paramagnetic than V³⁺ (t_{2g}²e_g⁰) and generally causes a larger Fermi contact shift. Similar results are also observed in ²³Na spectra.

Turning to the ²³Na NMR spectra, the signals of Na₄VP are recorded at 8, 10, and 12 kHz to distinguish the isotropic signals and spinning side bands denoted by asterisks (Fig. 4a). A typical ²³Na spectrum of Na₃VP is presented in Fig. 4b and c, which includes two significant signals at 83 and 18 ppm with very different magnitudes and a set of spinning side bands. As previously reported by Hu *et al.*,⁵² the oxidation state of vanadium centers coordinated with Na at Na1 sites is higher than those coordinated with Na at Na2 sites, leading to a more significant ²³Na shift at the Na2 site in contrast with that at the Na1 site. Similar paramagnetic effects have also been reported in LIBs on ⁷Li NMR shifts.⁵³ Based on these results, the signals at 83 and 18 ppm are assigned to the partially occupied Na2 site and fully filled Na1 site, respectively. Only Na at the Na2 site has electrochemical activity on account of the relatively weak bonding. The signal peaks shrink and slightly shift to the higher direction with the increasing sodium content of these samples (Fig. 4b). This phenomenon suggests a more ordered environment of Na, whereas more V²⁺ leads to a higher shift. We speculate that the shrinkage of peaks from Na₃VP to Na₄VP is due to a decrease in Na⁺ motion. Because of the partially occupied Na2 site, Na₃VP has more Na vacancies than the fully occupied Na₄VP. We further investigate Na⁺ diffusion kinetics by the CV technique at varied scan rates (Fig. S15 and S16†). The Na⁺ diffusion coefficients of Na₄VP decrease as we suspected, indicating the reduced mobility of sodium ions. The fitted spectra of samples with different sodiated degrees are presented in Fig. 4c and d and we conclude that the Na2/Na1 peak

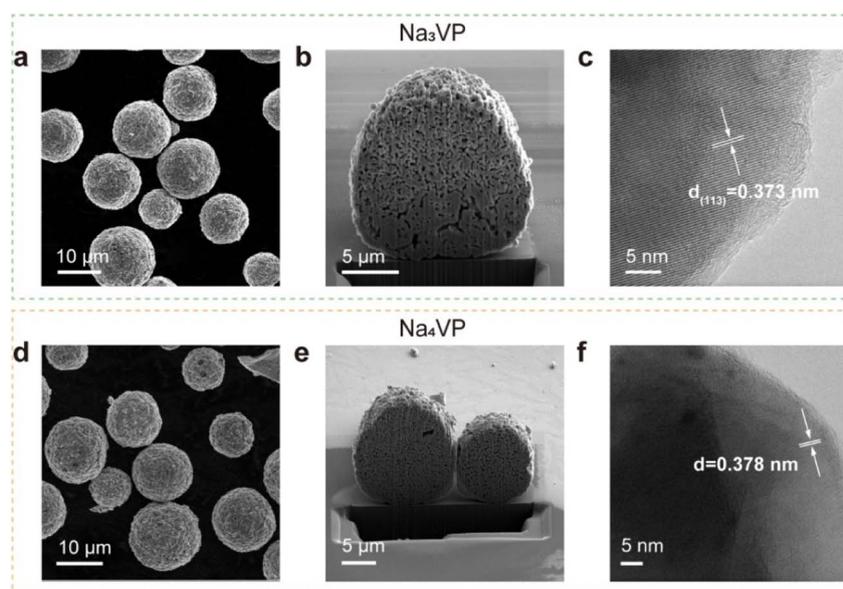


Fig. 3 (a–c) SEM image, cross section and HR-TEM image of Na₃VP, respectively. (d–f) SEM image, cross section and HR-TEM image of Na₄VP, respectively.



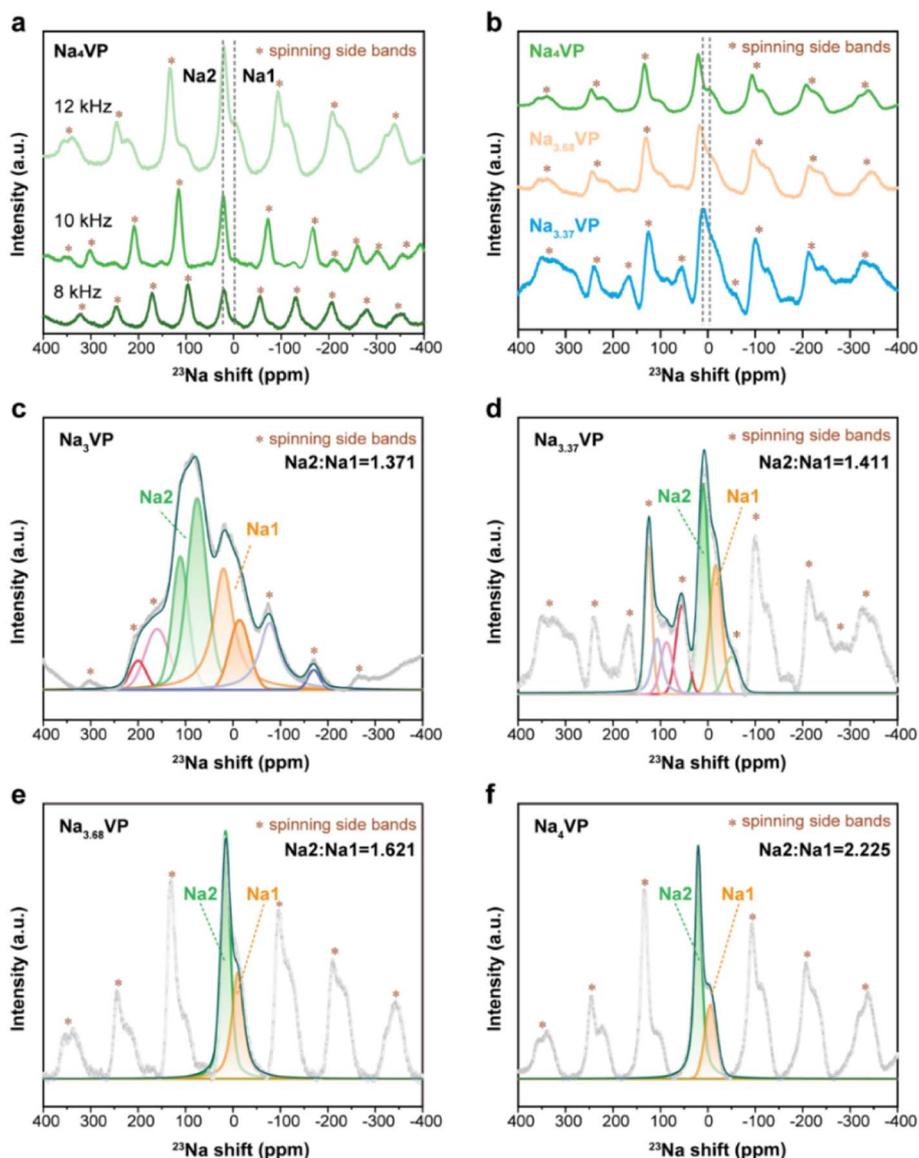


Fig. 4 (a) ^{23}Na ssNMR spectra of Na_4VP at 8, 10, and 12 kHz. (b) ^{23}Na ssNMR spectra of Na_xVP with different sodiation depths at 12 kHz. The fitted ^{23}Na ssNMR spectra of (c) Na_3VP , (d) $\text{Na}_{3.37}\text{VP}$, (e) $\text{Na}_{3.68}\text{VP}$, and (f) Na_4VP .

area ratio gradually increases from 1.371 to 2.225. In other words, excess sodium ions that Na_3VP receives from the PNZ- Na solution are preferentially accommodated in the Na2 site, making it close to completely occupied in the Na_4VP sample.

The electrochemical performance of the Na-enriched Na_4VP cathode

The sodium storage capability of the sodiated samples is investigated through galvanostatic charge–discharge tests in half-cells with a Na-metal anode. All of the Na_xVP electrodes with different treatment times manifest the same cycling stability at 0.5C for 100 cycles as pristine Na_3VP (Fig. S17[†]). Even extending the treatment time to 30 min, no discernible capacity decay is observed, suggesting that the electrochemical performance is completely retained after presodiation treatment. Our

presodiation strategy is not strict with processing time and is expected to achieve large-scale production. In the following discussion, we primarily focus on the comparison between pristine Na_3VP and the sodiated Na_4VP . Fig. 5a displays the CV curves of Na_4VP at a scan rate of 0.1 mV s^{-1} . The OCV of the Na_4VP sample is close to 1.6 V which is related to the average valence of V ions, in line with previous discussion. During the first cathodic scan, the first oxidation peak appears at around 1.7 V corresponding to the first distinct plateau in the charge profile of Na_4VP (Fig. 5b). These features, which are absent in Na_3VP electrodes, are because of the oxidation $\text{V}^{2+}/\text{V}^{3+}$ in the Na-enriched phase and the extraction of the pre-doped sodium ions. When all pre-doped sodium is removed, the Na-enriched Na_4VP entirely transforms into Na_3VP and then presents a typical voltage plateau of 3.4 V attributed to the reversible $\text{V}^{3+}/\text{V}^{4+}$ redox couple, validated by well-overlapped CV profiles. In



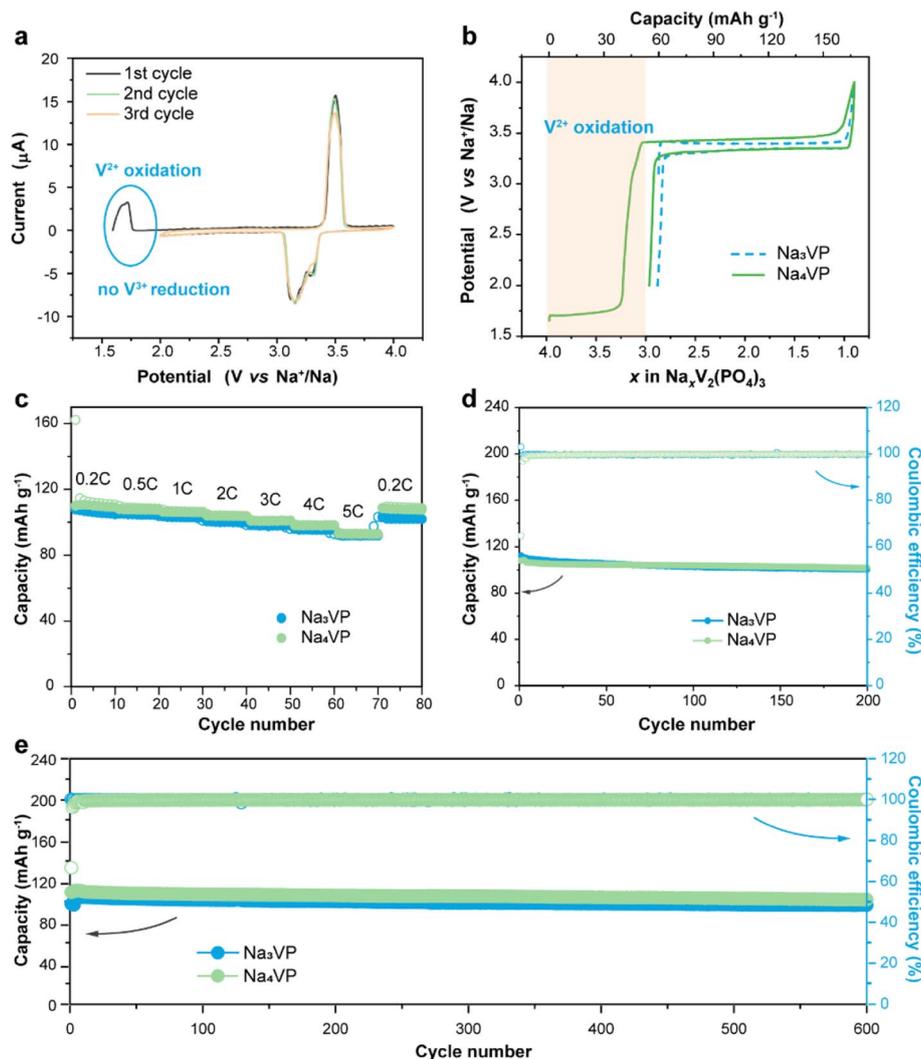


Fig. 5 (a) Cyclic voltammogram profiles of the Na₄VP cathode at a scan rate of 0.1 mV s⁻¹. (b) Initial charge–discharge curves of Na₃VP and Na₄VP electrodes between 2.0 and 4.0 V at 0.2C. (c) Rate performance, (d) cycling stability at 0.5C, and (e) long-term cycling stability at 1C of Na₃VP and Na₄VP electrodes.

brief, active sodium compensation during cycling is achieved by the irreversible oxidation of V²⁺, *viz.*, by controlling the voltage window. These results reconfirm that the preparation of the Na₄VP cathode to offer charge compensation and an additional Na reservoir has been successfully implemented. The rate capability and long-term cycling stability of the two samples are analysed as shown in Fig. 5c–e in a voltage window of 2.0–4.0 V (*vs.* Na⁺/Na). The Na₄VP electrode exhibits reversible desodiation capacities of 109.4, 107.5, 106.1, 101, 98.1, and 92.8 mA h g⁻¹ at rates of 0.2, 0.5, 1, 2, 3, 4, and 5C, respectively, slightly higher than that of the Na₃VP electrode. When the rate comes back to 0.2C, a reversible capacity of about 108 mA h g⁻¹ is achieved, suggesting the excellent reversibility of the Na₄VP electrode. The Na₄VP electrode also retains an outperformed cycling stability with high capacity retentions of 95.3% at 0.5C after 200 cycles and 91.8% at 1C after 600 cycles.

Na₄VP||HC full cells are assembled to demonstrate the application feasibility of Na₄VP as both the cathode and extra

Na reservoir in enhancing the energy density of SIBs, and a control cell Na₃VP||HC is also constructed for comparison. For the fabrication of full cells, the mass matching between the cathode and anode is based on their capacity balance, which is determined by the charge capacity of the cathode and the discharge capacity of the anode. As displayed in Fig. 6a and b, during the first charge, both the irreversible sodium storage sites of the anode and the formation of the SEI will consume considerable active sodium. More Na₃VP cathode is commonly to be matched with HC and in a Na-poor state in the following cycles, namely Na_{3-x}VP. As an alternative, the Na₄VP electrode used for full cells enables a 35.2 wt% reduction in cathode mass compared to the Na₃VP electrode and ensures that the cathode as Na₃VP without Na loss participates in subsequent cycles.

As shown in Fig. 6c, the Na₃VP||HC and Na₄VP||HC full cells deliver a charge capacity of 111 and 168 mA h g⁻¹, respectively, matching well with the values in half cell configuration. However, the reversible discharge capacity of the Na₃VP||HC full



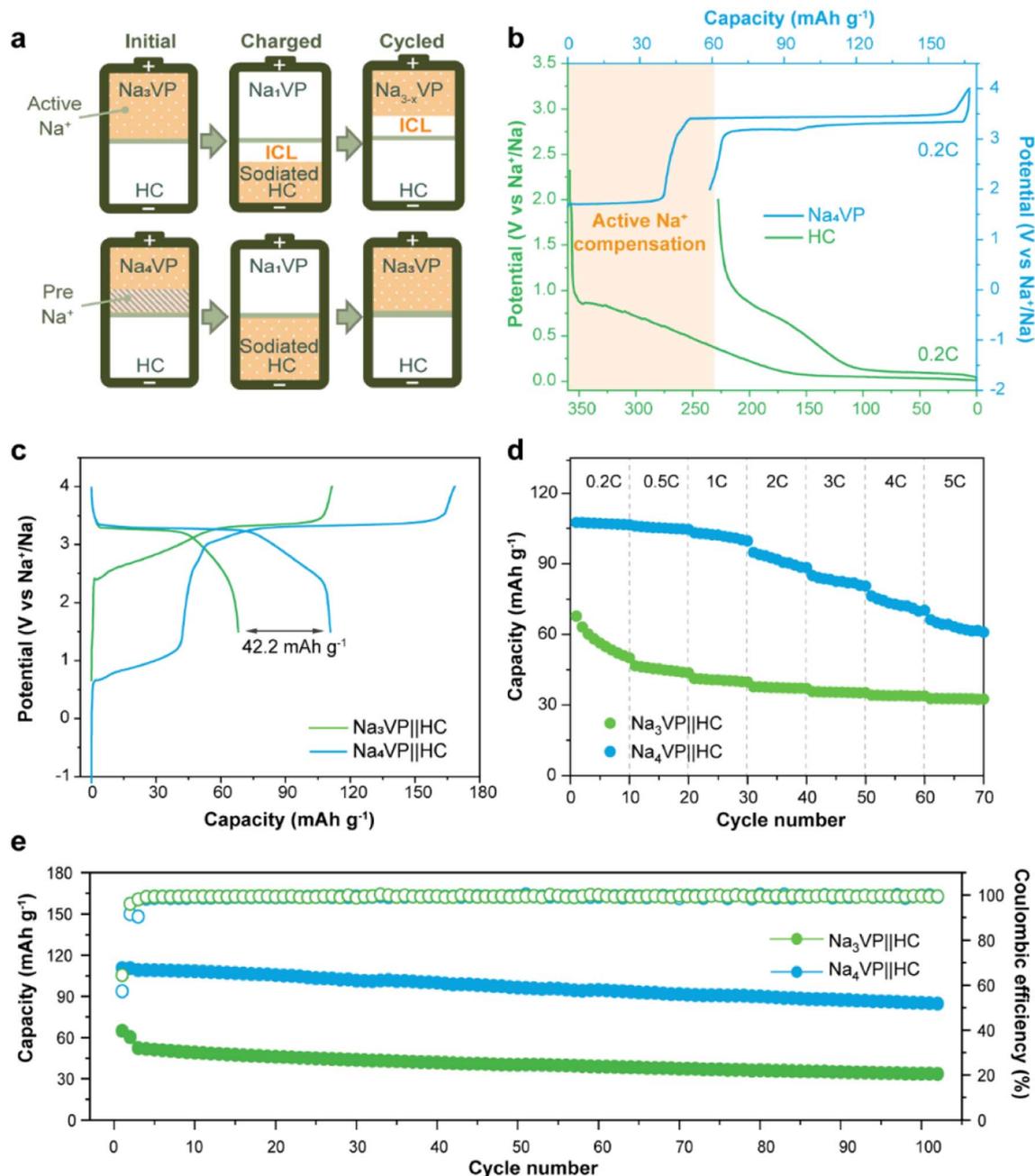


Fig. 6 (a) Schematic illustration of the presodiation process by the Na-enriched Na₄VP cathode in full cells. (b) Half-cell voltage profiles of the Na₄VP cathode and the HC anode at 0.2C. (c) Initial voltage profiles, (d) rate performance, and (e) cycling stability at 1C of Na₃VP||HC and Na₄VP||HC full cells.

cell is merely 66.6 mA h g⁻¹ with 39.3% capacity loss. This result can be attributed to the failed Na⁺ reinsertion on the cathode side in the absence of extra endogenous Na⁺. In comparison, the full cell with the Na₄VP cathode shows an initial discharge capacity of 109.7 mA h g⁻¹ and an ICE of 65%, which are equivalent to that of the Na₃VP half-cell, demonstrating extremely high utilization of active materials. Similar to other published results (Table S1[†]), the low ICE is ascribed to irreversible Na compensation during the initial charge cycle, but subsequent average coulombic efficiency is higher than that of

Na₃VP||HC without presodiation (Fig. S18[†]). The Na₄VP||HC full cell shows comparable cycling stability with the control Na₃VP||HC full cell, but shows markedly enhanced reversible capacity and excellent rate capability (Fig. 6d, e and S19). Fig. 6d demonstrates the rate capability of both Na₄VP||HC and Na₃VP||HC full cells from 0.2C to 5C. Na₄VP||HC displays a respectable rate performance benefiting from the presodated Na₄VP cathode. Discharge capacities of 107.5, 106, 103.2, 94.8, 84.9, 76.3, and 66.3 mA h g⁻¹ are delivered at current densities of 0.2, 0.5, 1, 2, 3, 4, and 5C, respectively. Even at a current



density of 5C, corresponding to a charging duration as short as 12 min, Na₄VP||HC still delivers 61.7% capacity retention of the discharge capacity at 0.2C. In contrast, Na₃VP||HC exhibits much lower reversible capacity due to active sodium loss. Overall, Na₄VP||HC exhibits an impressive discharge capacity of 111 mA h g⁻¹ with an ICE of 66%, and an energy density of 251.1 W h kg⁻¹ based on the total mass of cathode and anode active materials. This is far superior to those of the control cell using the pristine Na₃VP cathode (ICE of 60% and an energy density of 159.4 W h kg⁻¹). The Na₄VP||HC full cell shows a decent cycling stability of 78% capacity retention at 1C after 100 cycles and 57% capacity retention at 2C after 500 cycles (Fig. 6e and S19), which is comparable with that of the control Na₃VP||HC full cell, but shows markedly enhanced reversible capacity. Table S1 in the ESI† compares the recently published NVP-based full cells in terms of electrochemical performance and energy density. To the best of our knowledge, the energy density of PNZ-Na presodiated Na₄VP||HC is among the best of the reported values for NVP-based full cells. These discrepancies strongly suggest that our presodiation strategy can improve the overall energy density of SIBs while simultaneously maintaining decent cycling stability.

Conclusions

In summary, we have reported a facile and effective chemical presodiation approach to prepare a Na-enriched Na₄VP cathode under the guidance of the redox-potential-matching principle. The phenazine-Na reagent with a redox potential of 1.56 V vs. Na⁺/Na was employed to pre-dope active Na⁺ into the Na₃VP lattice for the Na₄VP cathode, successfully retaining the structure stability and electrochemical performance. Na₄VP initially serves as an excess Na reservoir and a cathode to eliminate the active Na⁺ loss on the anode side, allowing the full capacity of Na₃VP to be used for the subsequent cycles. Our presodiation strategy not only enables rapid preparation of Na₄VP and is expected to achieve batch production, but also doesn't bring about any external impurities after sodium compensation. The energy density of the Na₄VP||HC full cell is 251.1 W h kg⁻¹, much greater than that of the Na₃VP||HC full cell (159.4 W h kg⁻¹). Most importantly, this approach can be broadly applicable to different cathode materials and opens up new avenues for the design of high-energy SIBs.

Experimental section

Materials synthesis

All the reagents were used as received without further treatment. Commercial Na₃VP and hard carbon (BHC400) powders were purchased from Guangdong Canrd New Energy Technology Co., Ltd Phenazine, biphenyl, naphthalene and electron-grade N-methyl-2-pyrrolidone (NMP) were obtained from Aladdin Chem. Co., Ltd Anhydrous DME was purchased from JK Chemical. The phenazine-sodium (PNZ-Na) reagent was prepared by adding equal stoichiometric sodium metal and phenazine in DME and stirring for four hours in a glove box. Sodium-enriched Na₄VP was synthesized simply by immersing the Na₃VP powders or cathode

sheet in 0.05 M PNZ-Na solution for a certain time (0–120 s) and then rinsing it with DME to quench the reaction and remove the residue, followed by vacuum drying for 15 min.

Characterization

XRD patterns were collected on an X-ray diffractometer (Rigaku Smartlab SE) with Cu-Kα radiation at 40 kV. *In situ* XRD was performed on a live charge/discharge process to monitor the structure evolution of Na₃VP. Every single pattern was collected for 7 min in the scan range of 18–38°. The morphologies and fine phase structure of samples were obtained by SEM (SEM, Zeiss Merlin Compact) and HRTEM (JEM-2100). The valent information was obtained by XPS (ESCALAB250Xi) with Al Kα radiation. Raman spectra (Bruker MultiRam) were measured to investigate the variation of chemical bonds before and after persodiation. The chemical compositions of samples were determined by ICP-OES (Agilent 730) and EDS (X-Max 50). FIB was used to observe the cross section of active material particles on Tescan SOLARIS equipment. The ssNMR spectra were recorded with a Bruker AVANCE NEO 400 MHz NMR spectrometer.

Electrochemical measurements

The Na₃VP cathode sheet was fabricated by mixing Na₃VP, Super P and polyvinylidene fluoride at a mass ratio of 80 : 10 : 10 in NMP and further spreading the mixture on the aluminum foil with a doctor blade. After vacuum drying, the electrode was cut into discs with a diameter of 12 mm for use. For half cells, CR-2032-type coin cells were assembled by using pieces of Na foil as the counter and reference electrodes, Na₃VP or Na₄VP as the working electrode, and glass fiber as a separator in an argon-filled glove box with H₂O and O₂ contents less than 0.1 ppm. The electrolyte was 1 M NaClO₄ dissolved in ethylene carbonate/diethyl carbonate (1 : 1 v/v) solution with a 5% fluoroethylene carbonate (FEC) additive. Galvanostatic charge/discharge tests were conducted between 2.0 and 4.0 V vs. Na⁺/Na on a NEWARE CT4008 cycler battery cycler at room temperature. In particular, the nominal capacity of 1C is 117.6 mA g⁻¹. For full cells, all procedures were identical to those for half-cell assembly with the exception of using a hard carbon electrode as the anode. The cathode/anode capacity ratio was set as 1 : 1 and the voltage range was set as 1.5–4.0 V vs. Na⁺/Na. The CV measurement was executed on an Admiral electrochemical workstation at a scan rate of 0.1 mV s⁻¹. The energy density (*E*) of the full cell was calculated using the following formula:

$$E = Q \times V / (m_{\text{cathode}} + m_{\text{anode}}) \quad (6)$$

where *Q*, *V*, *m*_{cathode}, and *m*_{anode} indicate the discharge capacity and average discharge voltage and the mass of the cathode and anode, respectively.

Data availability

All the data supporting this study are included in the manuscript and the ESI.† Example data sets are available on request.



Author contributions

J. Q. and M. X. conceived the idea of the study. M. X. contributed to the design of the research and the experimental data analysis. F. Z. and Y. Z. conducted the XRD experiments. C. W. conducted the SEM experiments and related data analysis. X. Z. conducted the ssNMR experiments and the analysis of ssNMR results. X. A. and J. Q. supervised the work. The manuscript was written by M. X. and all authors participated in reviewing the manuscript.

Conflicts of interest

The authors declare no conflicts of interest.

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