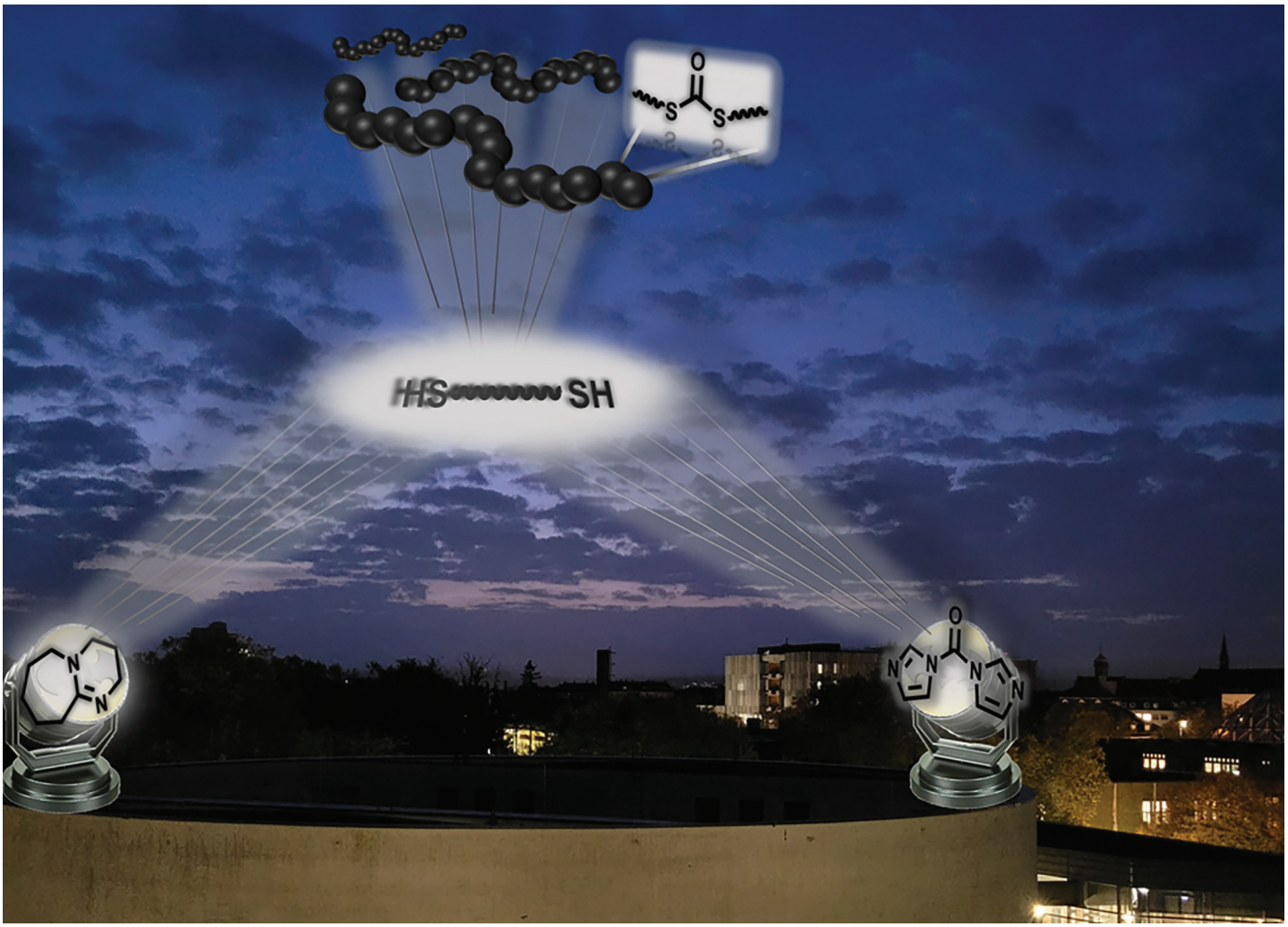

Highlighting research from Dr Hatice Mutlu and co-workers within the Soft Matter Synthesis Laboratory at Karlsruhe Institute of Technology, Germany.

Straightforward synthesis of aliphatic polydithiocarbonates from commercially available starting materials

A facile synthetic strategy towards functional and structurally diverse polydithiocarbonates, which reveal unexplored intrinsic emission properties, was developed under benign and mild reaction conditions by taking advantage of the versatile reactivity of 1,1′-carbonyldiimidazole (CDI) with dithiols in the presence of diazabicyclo[5.4.0]undec-7-ene (DBU), as an efficient base.

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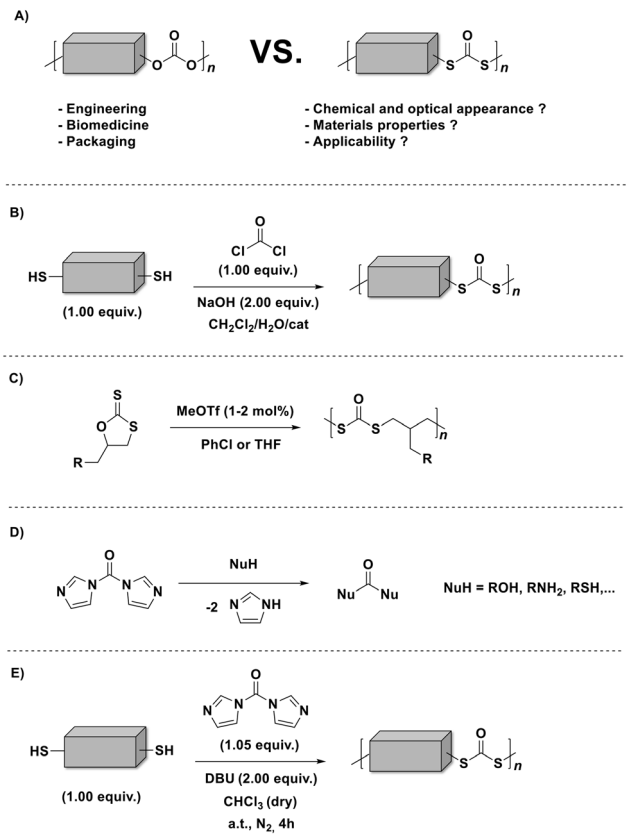
Straightforward synthesis of aliphatic polydithiocarbonates from commercially available starting materials†

Timo Sehn, Birgit Huber, Julian Fanelli and Hatice Mutlu *

Herein, a novel 1,1′-carbonyldiimidazole (CDI) mediated polymerization methodology that complements ROP and unlocks a greater synthetic window to less-recognized polydithiocarbonates is presented.

Sulfur containing polymers[1,2] are fascinating macromolecules that have received increasing attention due to their unique properties (*e.g.*, affinity to metals, improved refractive index, divergent redox and thermal characteristics, and enhanced chemical and biological properties). As a matter of fact, sulfur tethered macromolecules [particularly, polythiocarbonates, polydithiocarbonates (depicted in Scheme 1A) and polytrithiocarbonates] show higher susceptibility to light induced degradation (although they are recognized to be non-biodegradable),[3,4] compared to their oxygen counterparts (*i.e.*, polycarbonates). In this way, those polymers could even facilitate new light-induced chemical recycling methods as an answer to the ever-increasing amount of plastic pollution.[5] Indeed, several (degradable) polymeric scaffolds decorated with various sulfur-functionalities (*e.g.*, thioethers, thioesters, thiocarbonates, thiourethanes, thioamides, and thioureas among others) are now accessible.[1] Still, compared to the more conventional oxygen-rich polymers (such as aliphatic polycarbonates),[6] the toolbox of polymerizations to furnish sulfur-rich polymers (particularly, aliphatic polydithiocarbonates in which both of the ethereal oxygen atoms of the carbonate group are replaced by sulfur atoms) remains limited to polycondensation (Scheme 1B)[7] and the ring-opening polymerization (ROP) of cyclic sulfur-containing monomers (Scheme 1C).[8] The mentioned approaches are usually accompanied by unavoidable drawbacks involving the usage of toxic phosgene (or chloroformates) along with the tedious and low yield multiple synthetic steps of the cyclic monomers. Indeed, the absence of practical synthetic strategies that unlock a greater number of functional aliphatic polydithiocarbonates has most likely inhibited their broader use in appli-



Scheme 1 Schematic representation of: (A) the replacement of oxygen with sulfur atoms in polycarbonates, resulting in polydithiocarbonates; previously reported approaches for polydithiocarbonate synthesis: (B) phosgene-based step-growth and (C) ring-opening polymerizations; (D) The reactivity of **CDI** with diverse nucleophiles; (E) the currently investigated **CDI** mediated polymerization for polydithiocarbonate synthesis in the presence of DBU.

Soft Matter Synthesis Laboratory, Karlsruhe Institute of Technology (KIT), Hermann-von-Helmholtz-Platz 1, D-76344 Eggenstein-Leopoldshafen, Germany.
E-mail: hatice.mutlu@kit.edu

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cation-driven research. Although, aliphatic polydithiocarbonates could offer parallel properties and applications to those of the corresponding polycarbonates.

Aliphatic polydithiocarbonates, as analogues of polycarbonates, are expected to have improved thermal properties. In fact, a literature survey reveals[7,9–11] that higher melting temperatures have been reported for aliphatic polydithiocarbonates with respect to the oxygenated analogues.[12,13] Thus, in particular, poly (hexamethylene dithiocarbonate) has a melting temperature (T_m) about 60 °C higher than that of poly(hexamethylene carbonate), *i.e.*, 90 °C *vs.* 30 °C.[10] The latter behaviour could be attributed to the higher packing efficiency in the crystalline lattice when sulfur atoms are present in the polymer backbone.[14] In other words, poly(hexamethylene dithiocarbonate) is a material with a higher degree of crystallinity compared to poly(hexamethylene carbonate), respectively about 60% *vs.* 15%, and thus, reflects a faster crystallisation rate. Moreover, polymer chains decorated with –S–CO–S– groups are known to show higher flexibility and mobility compared to the chains containing –O–CO–O– groups.

Besides, the poly(hexamethylene dithiocarbonate) derivative shows a decomposition temperature (T_d) of ~320 °C (as it is discussed below) which is 40 °C higher compared to the oxygen analogue (~280 °C).[15]

1,1′-Carbonyldiimidazole (**CDI**, also known as Staab's reagent)[16] has been employed for various applications in organic chemistry,[17,18] and quite recently for polymer synthesis by virtue of its reactivity with various nucleophiles (carboxylic acids, alcohols, and amines) as shown in Scheme 1D.[19,20]

Moreover, **CDI** is safe, cheap (3 € per gram), and easy to handle, compared to the commonly used toxic phosgene or chloroformates. In fact, Malkoch[20] and colleagues reported the detailed phosgene-free synthesis of aliphatic polycarbonates with a broad range of properties. Surprisingly, literature survey has revealed that the synthesis of polydithiocarbonates in the presence of **CDI** is so far an unexplored research ground (Scheme 1E), despite the obvious advantages including easily available monomers (dithiols, which possess higher nucleophilicity compared to diols) and the relatively low toxicity of the imidazole by-product. Accordingly, the continuously increasing need for advanced materials prompted us to investigate a new synthetic approach with emphasis on operational practicality. In other words, we herein disclose a novel synthetic strategy that expands the window of aliphatic polydithiocarbonates that are synthesized from simple dithiol building blocks in the presence of a slight excess of **CDI** (*i.e.*, 1.05 equiv.). In fact, the main idea of utilizing the slight excess of **CDI** is to achieve tailor-made imidazole end-groups, which subsequently could be easily transformed into the corresponding methyl monothiocarbonate end-groups during the precipitation process in ice-cold methanol (the mechanism of the polymerization is depicted in Scheme S1†). In this way we ensure that no free-thiol end groups are present, which could give rise to polydisulfide formation. Given the fact that diazabicyclo[5.4.0]undec-7-ene (DBU) has been recognized as an effective organocatalyst for the **CDI**-mediated[21] nucleophilic reactions as well as in a variety of thiol-based[22] organic syntheses, we sought to examine the effect of DBU on the polymerization. Within the broad family of aliphatic polydithiocarbonates, the synthesis of poly(hexamethylene dithiocarbonate) is the most widely explored as this polymer distinguishes itself by virtue of its complex thermal behaviour.[10] Hence, to prove our hypothesis as a suitable strategy, an initial screening of the reaction conditions was performed using 1,6-hexane dithiol **M1** (shown in Fig. 1A) and **CDI** (Scheme 1E and Table S1†) in the presence of different equivalents of DBU (Fig. S1†) in 0.8 M chloroform to deliver polymer **P1**. In particular, by screening 2.00, 0.10 and 0.01 equivalents of DBU, the contribution of the base to the reaction and its impact on the degree of polymerization were investigated. Indeed, the efficiency of the base could be confirmed by using 2.00 equivalents of DBU resulting in polymer **P1** with an apparent number average molecular weight (M_n) of 11 500 g mol^{-1} and dispersity ($Đ$) of 2.1 in 4 hours at ambient temperature (Fig. 1B). Notably, the reaction was also effective using lower loadings of DBU (0.01 equivalent), however resulting in a lower M_n. The polymerization reaction which was carried out with no base (**P1.3** in Table S1†) showed that only a very low molecular weight polymer was formed regardless of the reaction time. These results underpin the non-catalytic nature of the base in this reaction. It is assumed that DBU is removed from the reaction mixture by scavenging the formed by-product imidazole to yield an ionic liquid (compare Scheme S1†),[23] thus reducing the concentration of the active base in the reaction solution. Nevertheless, the formation of an ionic liquid (IL) could be beneficial for the synthesis of the targeted polydithiocarbonates, as it has been recently reported that ionic liquids (ILs) can pave the way towards an eco-compatible route for synthesizing polycarbonates from sterically hindered alcohol derivatives (*i.e.*, isosorbide) and dimethyl carbonate.[24] The authors clearly emphasized that the cation and anion of ILs should synergistically activate the substrates. Indeed, the formation of the ionic liquid was also clearly observed with the aid of NMR analysis (particularly ^{13}C NMR), which was performed on the crude reaction mixture after 2 hours of the polymerization to deliver **P1** (Fig. S2 in the ESI†). The respective NMR data were compared with the ^{13}C NMR spectrum of the freshly prepared DBU–imidazole ionic liquid mixture in addition to the NMR data of pure DBU and pure imidazole (all data were collected in $CDCl_3$). Accordingly, a magnetic resonance at 162.4 ppm (compare Fig. S2†) that can be assigned to the protonated DBU species ($DBUH^+$) of the formed ionic liquid (which is consistent with the literature value, *i.e.*, 162.1 ppm)[25] was detected. Although, the literature survey[24] has emphasized that the cation and anion of ILs could synergistically activate the polymerization reaction, the stoichiometric ratio of DBU was increased, yielding the highest molecular weights, when 2.00 equiv. were employed. Crucially, the beneficial features of **CDI** include its low nucleophilicity and distinct neutral character, enabling the effective formation of polydithiocarbonates without the need for the continuous removal of the imidazole by-product (without the need for any additional setup).

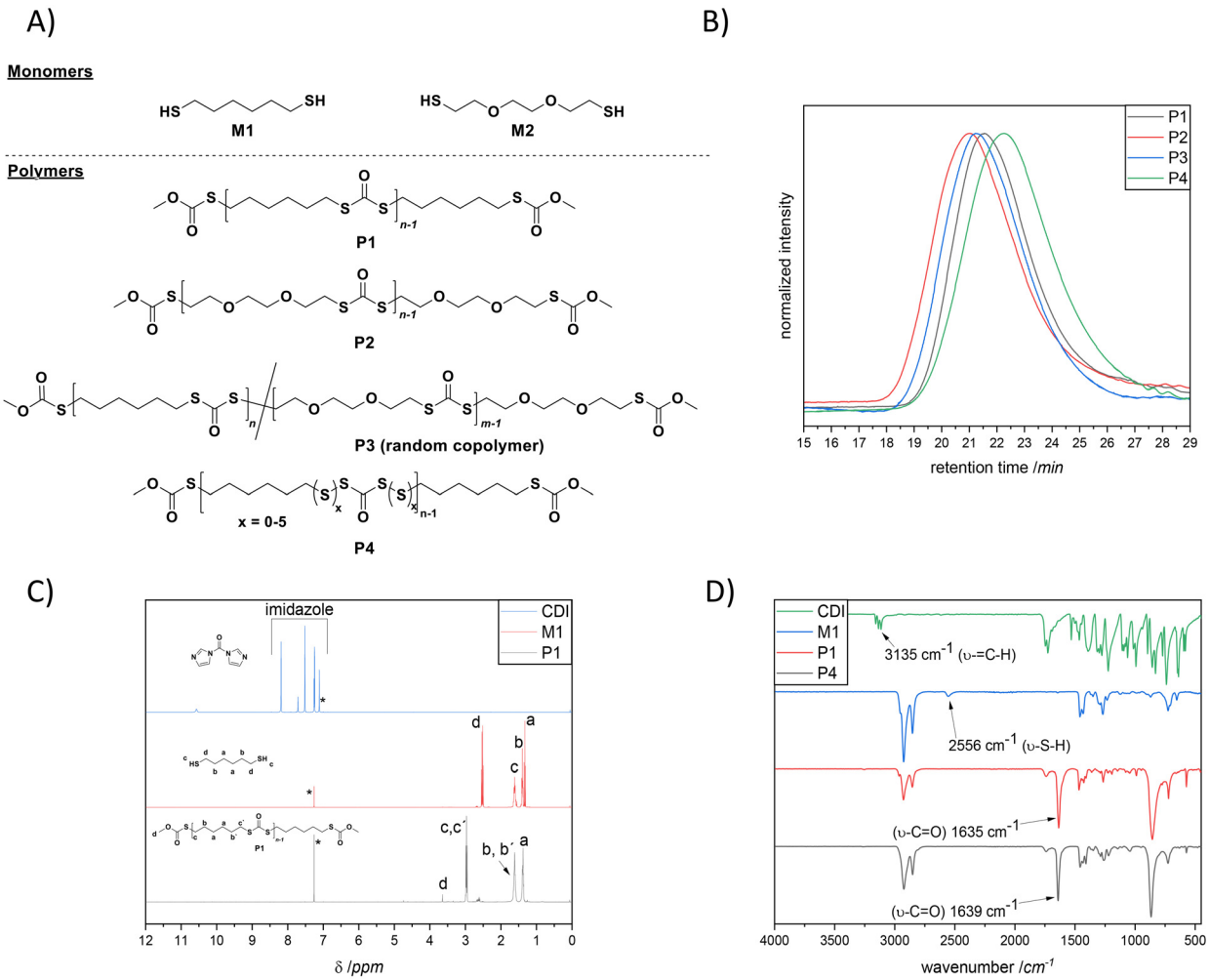

Fig. 1 (A) Chemical structures of monomers **M1** and **M2** in addition to the synthesized (co)polymers **P1–P4**. (B) SEC traces of step growth polymers **P1** (M_n = 11 500 g mol^{-1}, $Đ$ = 2.1, black line), **P2** (M_n = 14 700 g mol^{-1}, $Đ$ = 2.4, red line), **P3** (M_n = 14 000 g mol^{-1}, $Đ$ = 2.1, blue line) and **P4** (M_n = 6500 g mol^{-1}, $Đ$ = 2.5, green line) in THF + 0.2% w/v BHT. (C) ^{1}H NMR spectra (400 MHz, $CDCl_3$, 298 K) of **CDI** (up, blue line), **M1** (middle, red line) and **P1** (bottom, black line). (D) ATR-IR spectra of **CDI** (green line), **M1** (blue line), **P1** (red line) and **P4** (black line).

The polymerization of **P1** was monitored by Size Exclusion Chromatography (SEC) of reaction aliquots withdrawn at 30, 90, 120 and 240 min. In fact, the results depicted in Fig. S3† together with the mechanism postulated in Scheme S1† suggest a step-growth polymerization mechanism. Nevertheless, further studies (which are currently performed in our laboratories) are necessary in order to conclude the latter statement.

The successful incorporation of dithiocarbonate units into the polymer backbone was further demonstrated by NMR (1D and 2D), ATR-IR and UV-Vis analyses. The ^{1}H NMR spectrum and peak assignments of **P1** are shown in Fig. 1C, where explicitly the appearance of magnetic resonance corresponding to the –CH_2– groups at the α-position to the dithiocarbonate, *i.e.*, at 2.97 ppm, was observed. Moreover, the complete disappearance of the magnetic resonances of terminal –SH at 1.33 ppm and imidazole protons (appearing from 7.11 to 8.19 ppm) also confirmed the successful polymerization. In a complementary way, a new magnetic resonance that arose at 190 ppm was detected in ^{13}C NMR (Fig. S4†), attributable to the carbonyl carbon of the dithiocarbonate moiety in **P1** (which is also consistent with the literature).[26] Furthermore, 2D NMR, *i.e.* ^{1}H–^{1}H correlation spectroscopy (COSY) and heteronuclear single quantum coherence (HSQC) spectroscopy (Fig. S5 and S6 in the ESI†), confirmed in a supplementary manner the formation of the targeted polymer structure. Besides, the attenuated total reflectance infrared (ATR-IR) spectrum of **P1** shown in Fig. 1D revealed the characteristic C=O stretching vibration band of dithiocarbonate moieties at ~1635 cm^{-1}, while neither the =C–H stretch vibration band of imidazole nor the S–H stretch vibration band of thiol groups could be observed (appearing in the ranges of 3000 to 3150 cm^{-1} and 2550 to 2598 cm^{-1}, respectively). Moreover, a band at 871 cm^{-1} due to the C–S stretching was also detectable. Importantly, NMR and ATR-IR analyses affirmed that the resulting polymers contain only the expected polydithiocarbonate segments. Neither polydisulfide and polymonothiocarbo-

nate segments nor poly(*S*-dithiocarbonate) derivatives (polymers containing the oxythiocarbonylthio or xanthate ester) were observed. The latter products are attributed to the side reactions occurring during the ROP of cyclic monomers. Furthermore, the experimentally obtained elemental analysis values for **P1** (C: 49.7%, H: 6.81%, S: 37.9% in Table S2†) were in good agreement with theoretical ones, reaffirming the fact that the dithiocarbonate units were installed.

From a commercial point-of-view, a relevant advantage of the proposed strategy is that it allowed the synthesis of copolymers. The latter was manifested by the efficient random copolymerization of **M1** with **M2** (shown in Fig. 1A). Note to the reader: **M2** was the choice of dithiol because such a hydrophilic ethylene glycol segment could be introduced; hence the final copolymer was anticipated to show unique thermal properties. Prior to the copolymerization of **M1** with **M2**, **M2** was also homopolymerized with **CDI** under the optimized conditions (depicted in the ESI†), to ensure that chloroform, which has been utilized for the polymerization of **M1**, was also a suitable solvent for the polymerization of this slightly hydrophilic dithiol derivative. Indeed, the SEC results shown in Fig. 1B affirmed the successful homopolymerization and copolymerization attempts by delivering polymers with the respective M_n of 14 700 g mol^{-1} ($Đ$ = 2.4, red line, **P2**), and M_n = 14 000 g mol^{-1} ($Đ$ = 2.1, blue line, **P3**). The detailed characterization of the homopolymer **P2** and the copolymer **P3** is additionally provided in the ESI (Fig. S7 to S12†). Nevertheless, it is of crucial importance to emphasize that the incorporation of both dithiol monomers in **P3** was confirmed *via* in-depth spectroscopic analysis, *e.g.*, 1D and 2D-NMR in addition to ATR-IR. Indeed, the magnetic resonances at 2.95 and 3.18 ppm (^{1}H NMR, Fig. 2A) in addition to 30.2 and 30.5 ppm (^{13}C NMR, Fig. S10†) respectively affirmed the postulated chemical structure. On the one hand, the integral ratio of the magnetic resonances corresponding to c and d suggests 50 : 50 random copolymer formation (Fig. 2A). On the other hand, one could not observe any distinguishable pattern in the comparative IR spectra of **P2** compared with **P3** (Fig. 2B). Nevertheless, a complementary elemental analysis affirmed the copolymer composition (C: 48.5%, H: 6.77%, S: 35.4%).

Additionally, the efficiency of the polymerization approach was demonstrated by the synthesis of a polydithiocarbonate possessing a higher sulfur content, *i.e.*, **P4** (Fig. 1B and D, and Table S3 in addition to Scheme S2†) which was realized by introducing 0.5 equivalent of elemental sulfur, an abundant side product of the petrochemical industry, into the polymerization media. It is important to note that superbases, such as DBU, are known to catalyse diverse reactions of elemental sulfur at ambient temperature.[27] Subsequently, in addition to SEC, NMR and ATR-IR analyses (depicted in Fig. 13 to 17 in the ESI†), the elemental analysis of the isolated polymer, **P4**, showed a 4.4% higher sulfur content (42.3%, Table S3 in the ESI†) compared to **P1** (37.9%, Table S2 in the ESI†), hence underpinning the efficient incorporation of additional sulfur atoms into the polymer backbone. It is important to mention that the sulfur content in the repeating unit of any polymeric structure is thermodynamically limited to a maximum of 9 sulfur atoms before the homopolymerization of elemental sulfur starts to compete with depolymerization. The latter is observed as the average bond dissociation energy of the –S–S– bond is expected to decrease to a certain value, resulting in an accelerated sulfur exchange reaction.[28] Hence, at the most, six sulfur atoms can be integrated. Indeed, the multiplet in the region between 2.62 and 2.50 ppm associated with the hydrogen atom at the α-carbon adjacent to the disulfide bond was shifted to the following ranges: 3.03–2.93, 2.93–2.87 and 2.85–2.79 ppm, respectively, in such a way that each series of multiplets is assigned to the respective sulfur rank, hence nicely showing the high sulfur content of the polymer. In fact, this shift in the resonances of the disulfide bond in NMR analyses has been noted for polysulfides whose α-carbon protons exhibit multiplets, as in the current case.[28]

The synthesized polydithiocarbonate derivatives **P1–P4** were investigated regarding their thermal characteristics by thermogravimetric analysis (TGA) and differential scanning calorime-

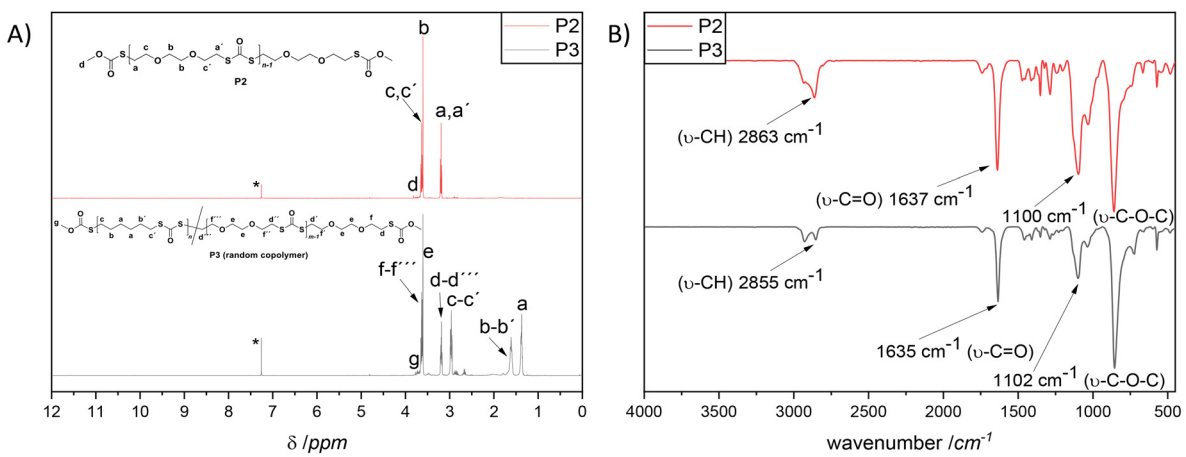

Fig. 2 (A) ^{1}H NMR spectra (500 MHz, $CDCl_3$, 298 K) of **P2** (up, red line) and **P3** (bottom, black line), respectively, homopolymer of **M2** in addition to a copolymer of **M1** and **M2**. (B) ATR-IR spectra of **P2** (up, red line) and **P3** (bottom, black line).

try (DSC). On the one hand, the decomposition temperature (T_d), defined as the temperature at which 5% weight loss takes place, was observed for **P1** at 321 °C (Fig. 3A), because of the highly crystalline behaviour. On the other hand, **P2–P4** showed lower thermal stability, *i.e.*, 258 °C < T_d < 280 °C. Importantly, all (co)polymers displayed a single step thermal degradation

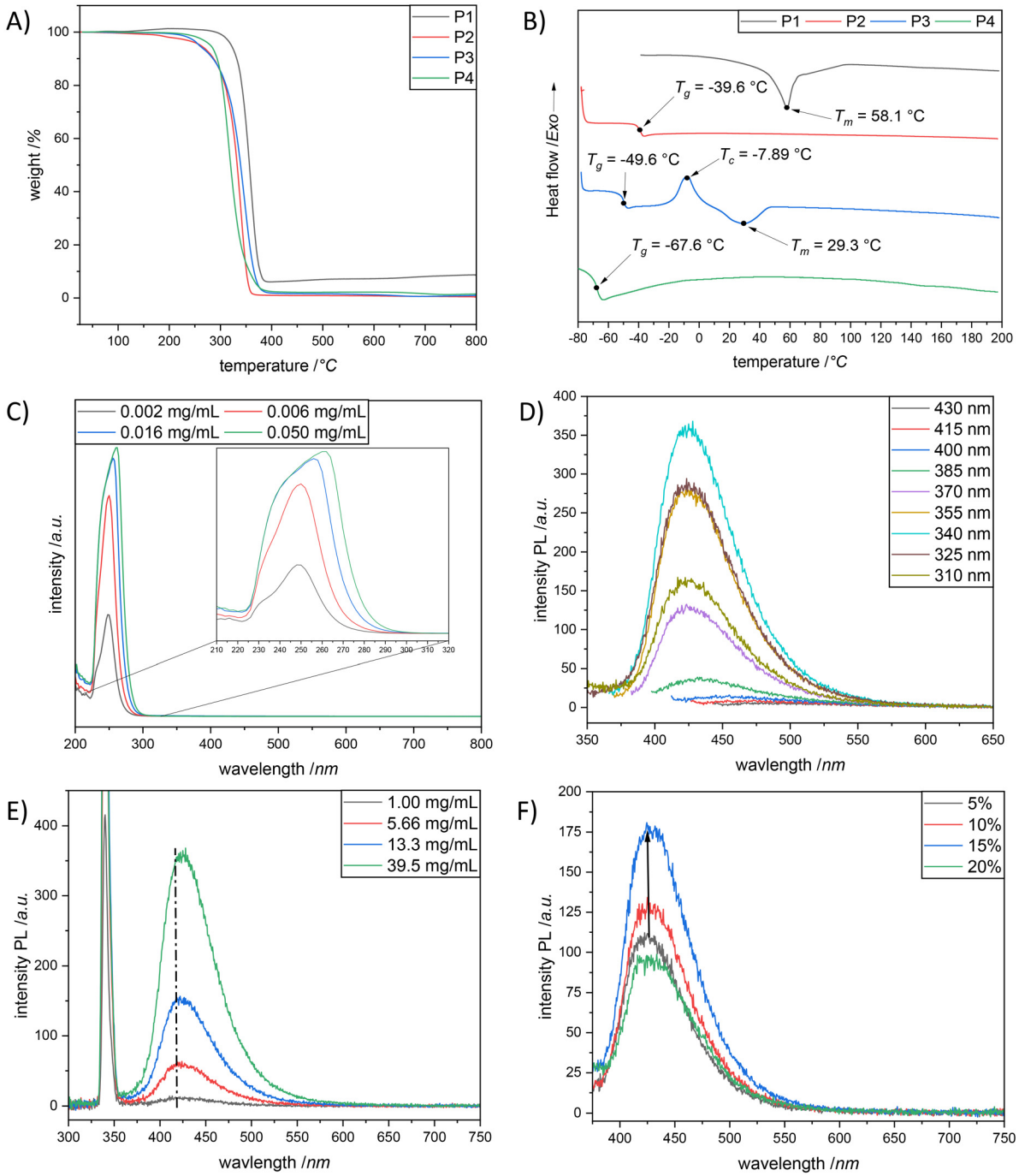

Fig. 3 (A) Thermogravimetric analysis (TGA) of **P1** (black line), **P2** (red line), **P3** (blue line) and **P4** (green line) from 25 to 800 °C with a heating rate of 10 K min^{-1} under a nitrogen flow. (B) Differential scanning calorimetry (DSC) studies (second heating run) of **P1** (black line), **P2** (red line), **P3** (blue line) and **P4** (green line) from −80 °C to 200 °C with a heating rate of 10 K min^{-1} under a nitrogen flow (as for **P1**, measurement was from −40 °C). (C) Absorption traces of **P2** at different concentrations (black: 0.002 mg mL^{-1}, red: 0.006 mg mL^{-1}, blue: 0.016 mg mL^{-1} and green: 0.050 mg mL^{-1}) in DCM (298 K). (D) Emission spectra of **P2** at various excitation wavelengths (from 310 nm to 430 nm) in DCM (c = 39.5 mg mL^{-1}). (E) Emission spectra of **P2** at different concentrations (black: 1.00 mg mL^{-1}, red: 5.66 mg mL^{-1}, blue: 13.3 mg mL^{-1} and green: 39.5 mg mL^{-1}) in DCM (298 K) at λ_{exc} = 340 nm. (F) Emission spectra of **P2** in DCM/water mixtures with an increasing water ratio (black: 5%, red: 10%, blue: 15% and green: 20%) at 298 K; the arrow indicates the slight shift of the maxima of each emission spectrum.

with no char residue (aside **P1**). Complementary to TGA, the DSC analysis confirmed the crystalline behaviour of **P1**. In other words, the second heating scan (β = 10 °C min^{-1}) of **P1** was characterised by the presence of a well distinct endothermic peak located at 58 °C (Fig. 3B, black line). The DSC study of **P2** indicated that the polymer was in the amorphous state, being characterized by a glass transition temperature, T_g, value of −40 °C. Whereas, copolymer **P3** exhibited a crystalline polymer behaviour and possesses amorphous domains, reflected with a glass transition at −50 °C (resulting from the hydrophilic monomer corporation, **M2**), followed by crystallization at −8 °C and the respective melting at 29 °C (the latter two underpinning the presence of a crystalline domain and hence the incorporation of monomer **M1** (Fig. 3B, blue line)). The high sulfur content counterpart of the crystalline polymer **P1**, *i.e.*, polymer **P4**, showed an amorphous state with a T_g of −67 °C (Fig. 3B, green line). Indeed, this is consistent with the results obtained by Fitch and Helgeson, who have shown that the increase in the sulfur content introduced into a polymer results in a lowering of the T_g.[29] Moreover, the domination of the amorphous phase and the lack of crystallinity in **P4** confirm the random arrangement of the repeating units of the high sulfur macromolecules.

The discovery of intrinsic fluorescence in non-aromatic materials (such as proteins, peptides, and rice amongst others,[30] which was later termed a clustering-triggered emission (CTE)[31] mechanism) has advanced the current fluorescence paradigm. In particular, because this discovery not only extends the traditional principles and establishes new theories,[32–36] but also improves the virtue of biocompatibility and biodegradability, aliphatic fluorescent polymers have shown more privileges in the biomedical and biological fields over the existing fluorescent chromophore containing aromatic moieties.[37,38] The so-called non-conventional luminescent compounds generally comprise various isolated chromophores, such as electron-rich heteroatoms (*i.e.* N, O, and P),[39,40] hydroxyl (OH), amino (NH_2), carbonyl (C═O), cyano (CN),[30] carboxyl (COOH),[36] amide (CONH),[41] and anhydride[42] groups and carbon–carbon double bonds (C═C).[43] Accordingly, the formation of clusters of those isolated chromophores with π and/or n electrons is essential for the emission, which in turn establishes the CTE mechanism. While sulfur atom (S) has been proved to facilitate fluorescence due to larger atomic radius,[44] the non-conventional luminophores (small- or macro-molecules) involving the S moiety have rarely been reported.[45,46] As a matter of fact, we were interested if the herein synthesized polymers would be emissive under UV irradiation in aggregates, despite the absence of remarkable conjugations. In fact, all (co)polymers appeared as "see-through" transparent solids (or highly viscous liquids) under natural light, whereas a blue emission was observed under 365 nm UV light (as shown in Fig. S18†). It is important to mention that recently non-conjugated (hyperbranched) polycarbonate derivatives were reported to exhibit bright blue photoluminescence without any additional treatment under 365 nm UV light.[47–49] The authors attributed the fluorescence emission to the cluster-induced emission (CIE) mechanism which resulted from the clustering of the locked carbonyl groups, although they could not postulate a clear mechanism to these non-aromatic fluorescent materials. Inspired by the above research, the absorbance and photoluminescence properties of **P2** in dichloromethane (DCM) were carefully studied, as preliminary screening of all polymers revealed that the strongest emission was observed for **P2** (compare Fig. 3D and Fig. S23, S25 and S26 in the ESI†). As depicted in Fig. 3C, two absorption bands, centered at ∼ 232 nm and 248 nm, were detected, respectively, when the measurement was performed in a low concentrated solution, *i.e.*, 0.002 mg mL^{-1} **P2**. Eventually, the absorption band appearing at 232 nm can be associated with the π–π* electron transition of dithiocarbonate functional groups, whereas the absorption peak arising around 248 nm can be assigned to the n–π* electron transitions of dithiocarbonate moieties (that is common for molecules containing a C═O group in *S,S*-dialkyl-dithiocarbonates). Importantly, this information is in accordance with the literature.[50] At this point it is also noteworthy that the intensities of the absorption bands rise gradually while increasing the solution concentration of **P2**, accompanied by a small bathochromic shift. The latter phenomenon was presumably due to the aggregation of polymer chains at the increased solution concentrations of **P2**, which respectively resulted in dithiocarbonate (–S–C(O)–S–) cluster formation (Fig. S24 in the ESI†). Indeed, overlapping electron clouds and delocalized electrons in the dithiocarbonate clusters ensure an extensive electron delocalization system, which is responsible for the red shift of the absorption bands (Fig. 3E). To further characterize the photoluminescence properties of **P2** in DCM, a series of fluorescence emission spectra at various excitation wavelengths (310 nm to 430 nm) were also recorded (Fig. 3D). As expected, the fluorescence intensity of the polymer solution of **P2** in DCM, at the constant concentration of 39.5 mg mL^{-1}, significantly altered with the increase of the excitation wavelength from 310 nm to 430 nm; a maximum emission wavelength of 423 nm was detected at λ_{exc} = 340 nm (Fig. 3D). Additionally, a significant increase of the fluorescence emission intensity was observed for **P2**, while increasing the polymer concentration from 1.00 mg mL^{-1} to 39.5 mg mL^{-1} (Fig. 3E). The slight red shift of the emission band located at 416 nm to a higher wavelength of 423 nm (Fig. 3E) was also detected with the increase of the emission intensity as a result of the enhanced formation of rigid dithiocarbonate clusters at elevated polymer concentrations, which in turn underpins a possible cluster induced emission mechanism.[48] Because the dithiocarbonate functional groups are structural analogues of their oxygen analogues, *i.e.*, carbonate moieties, it was assumed that the cluster formation of dithiocarbonates is also triggered by space and n–π* electronic interactions as it was previously reported for carbonate moieties.[48] To further investigate the fluorescence properties of **P2**, the influence of incorporating different water fraction ratios into the solution of **P2** (dissolved in THF) was examined. Unsurprisingly, keeping the fact in mind that **P2** was not soluble in water, the fluorescence emis-

sion intensity of the corresponding polymer solution increased with elevated water fractions according to the typical aggregation-enhanced emission (AEE) effect (Fig. 3F).[48] The maximum fluorescence emission intensity of the polymer solution of **P2** was achieved with a water fraction of 15%. However, when the water fraction reached 20%, the polymer, *i.e.*, **P2**, precipitated, and the mixture started to become opaque. Consequently, the fluorescence emission intensity of the polymer solution faded at a water content of 20%. Hence, the experiment revealed in a different way that the macromolecular mechanics, *i.e.*, AEE effect, gives rise to the unexpected fluorescence character of the aliphatic polydithiocarbonates, specifically **P2** in our case.

Last but not least, in order to validate the abovementioned results and get more insights into the emission mechanism, a model compound, **MC**, mimicking the repeating unit of **P2** (Fig. 4A) was synthesized (Sections B6 and B7, and Fig. S25–S28 in the ESI†). The respective absorbance and emission data of this compound in DCM were recorded and compared with the optical properties of the starting compound 3,6-dioxa-1,8-octanedithiol in addition to the intermediate product 3,6-dioxa-1,8-octanedithio-*bis*-carbonylimidazole (Fig. 4B and C). The spectral patterns of these compounds (particularly, 3,6-dioxa-1,8-octanedithio-*bis*-carbonylimidazole and **MC** at 10 mg mL^{-1}) were similar to each other with a strong absorption band around 284 nm (with an obvious red-shift of 50 nm in comparison with that of 3,6-dioxa-1,8-octanedithiol) suggesting increased intramolecular cluster triggered interaction plausibly owing to the clusterization of the dithiocarbonyl moiety. Moreover, a shoulder around 370 nm was observed for 3,6-dioxa-1,8-octanedithio-*bis*-carbonylimidazole as a result of the n–π* electron transitions of the imidazole moieties.[51] Meanwhile, as a maximum emission wavelength of 423 nm was detected at λ_{exc} = 340 nm for **P2**, 3,6-dioxa-1,8-octanedithio-*bis*-carbonylimidazole and **MC** were subjected to analogue analysis; a slightly lower maximum emission wavelength of 351 nm was recorded at λ_{exc} = 340 nm. The intensity of 3,6-dioxa-1,8-octanedithio-*bis*-carbonylimidazole was slightly higher in comparison with that of **MC** (185 *vs.* 140 a.u.), which can be attributed to the presence of the imidazole and thiourethane units. In contrast, no emission was observed for 3,6-dioxa-1,8-octanedithiol. In a similar manner to **P2**, a significant increase of the fluorescence emission intensity was observed for **MC** while increasing the compound concentration from 1.00 mg mL^{-1} to 50 mg mL^{-1} (Fig. 4D). The slight red shift of the emission band located at 404 (5.00 mg mL^{-1}) nm to a slightly higher wavelength of 412 nm (50 mg mL^{-1}) was also in accordance with the results observed for **P2**.

On the one hand, these results may show that the dithiocarbonate unit could be both a CTEgen and an AIEgen, demonstrating the important role of clusterization in the manipulation of the photophysical behaviour of a molecular unit either in solution or in an aggregated state. On the other hand, to further

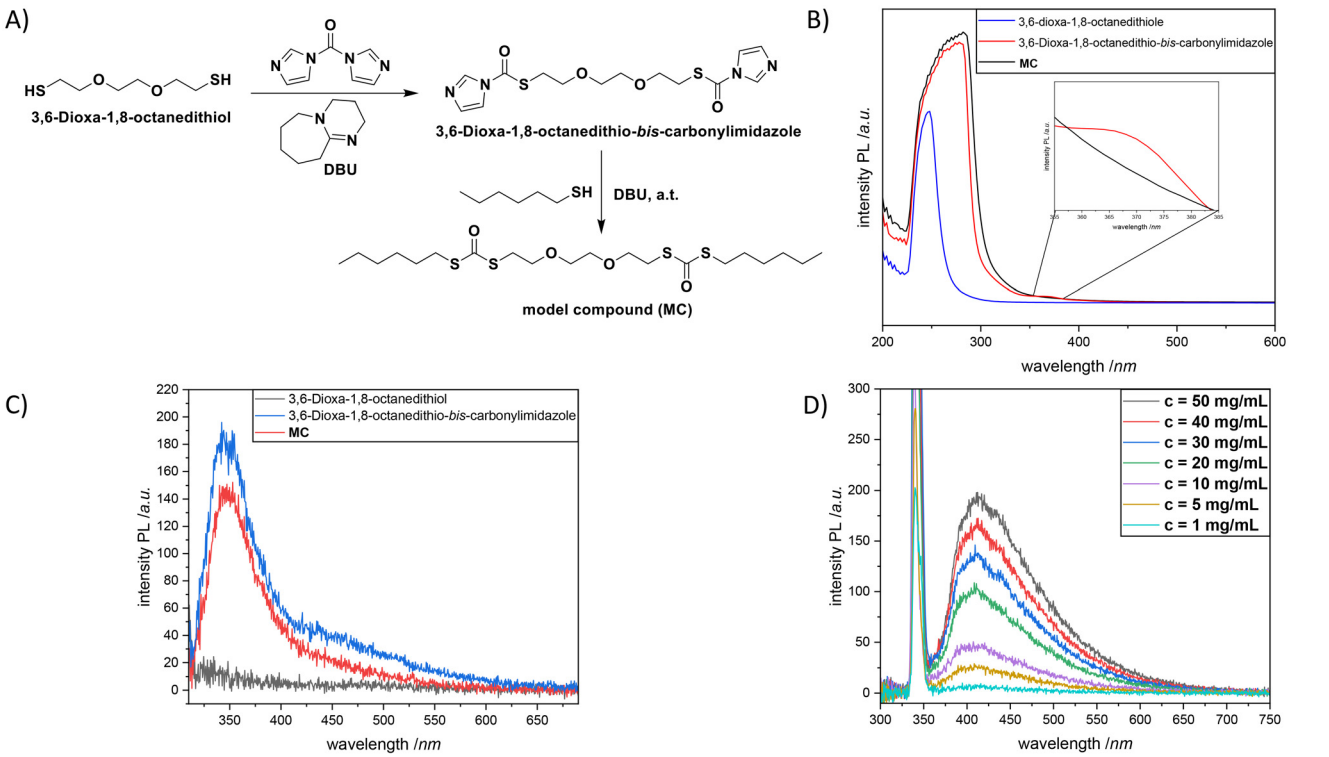

Fig. 4 (A) The synthesis of a model compound, **MC**, mimicking the repeating unit of **P2**. (B) Absorption traces of **MC**, 3,6-dioxa-1,8-octanedithiol and the intermediate product 3,6-dioxa-1,8-octanedithio-*bis*-carbonylimidazole in DCM (298 K). (C) Emission spectra of **MC**, 3,6-dioxa-1,8-octanedithiol and the intermediate product 3,6-dioxa-1,8-octanedithio-*bis*-carbonylimidazole in DCM (10 mg mL^{-1}) at λ_{exc} = 340 nm. (D) Emission spectra of **MC** at different concentrations (from 1.00 mg mL^{-1} to 50.0 mg mL^{-1}) in DCM (298 K) at λ_{exc} = 340.

understand the fluorescence behaviour of those classes of polymers and the respective small molecules at the molecular level, density functional theory (DFT) calculations are essential.

Conclusions

In conclusion, we have presented a facile synthetic strategy to obtain polydithiocarbonates (particularly poly(hexamethylene dithiocarbonate) and its copolymers) under benign and mild reaction conditions by taking advantage of the versatile reactivity of 1,1′-carbonyldiimidazole, CDI, with dithiols in the presence of diazabicyclo[5.4.0]undec-7-ene, DBU, as an efficient base. Crucially, the beneficial features of CDI include its low nucleophilicity and distinct neutral character, enabling the effective formation of polydithiocarbonates without the need for continuous removal of the imidazole by-product. The polymerization followed a traditional step-growth mechanism which was favored over depolymerization, and it is postulated that this approach should serve as a complement to the existing methodologies by combining the operational benefits of the ROP and scalability of traditional step-growth polymerization protocols. Indeed, the methodology facilitates the fabrication of previously inaccessible functionally and structurally diverse dithiocarbonate-decorated macromolecules, which exhibit unexplored photophysical properties. It is important to mention that studies on the degradation properties of those polymers are currently investigated. Thus, we truly believe that this approach will play a crucial role in the future availability of advanced polydithiocarbonates, and their use as materials in the technical and biomedical engineering fields, particularly on account of their unexplored intrinsic emission properties (which still need to be investigated in detail).

Conflicts of interest

There are no conflicts to declare.

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