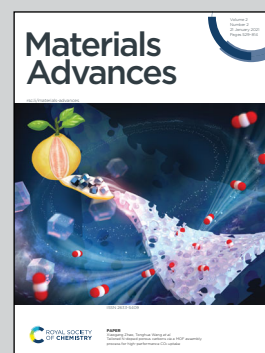


Showcasing research from Professor Junwang Tang's laboratory, Department of Chemical Engineering, University College London, London, United Kingdom.

Progress and challenges in photocatalytic ammonia synthesis

The group of Solar Energy and Advanced Materials, led by Prof. Junwang Tang focuses on photocatalytic small molecule activation (including CH₄, N₂, H₂O, C₆H₆ and CO₂) to high-value chemicals/fuels and microwave catalysis (e.g. plastic chemical recycling), together with microwave-intensified chemical processes. In parallel, the group has state-of-the-art spectroscopies to investigate the underlying charge dynamics and kinetics. This review puts forward a few critical issues existing in nitrogen photoreduction, along with a comprehensive discussion of strategies for improvement of NH₃ yield, aiming to make substantial contribution towards advancing this field.

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Progress and challenges in photocatalytic ammonia synthesis

Qing Han,[†] Haimiao Jiao,[†] Lunqiao Xiong^b and Junwang Tang^{*b}

Photocatalytic ammonia (NH₃) synthesis from N₂ and water driven by solar energy is a sustainable and environmentally friendly technology, which has gained considerable attention in recent years. In this review, the recent development in the fundamental understanding of photocatalytic NH₃ synthesis and the methods of precise NH₃ detection are summarized. More importantly the strategy for surface engineering and interface engineering of photocatalysts toward photocatalytic NH₃ production has been thoroughly analyzed with the aim to stimulate critical thinking about the effective methodology for catalyst modification instead of exploring new materials. At the end the challenges and a few concerns are raised from the current reports and future perspectives in this research field are discussed targeting to clarify the reliability and reproducibility of the photochemical process and to direct the future research direction, such as flow reactor design and in-depth understanding of the underlying reaction pathway.

Introduction

Ammonia (NH₃) has been regarded as one of the most important chemical products as feedstock for fertilizers and for various chemicals.^{1,2} The production of NH₃ plays a prominent role in global economy with an annual yield of more than 200 million tons.³ The industrial synthesis of NH₃ is dominated

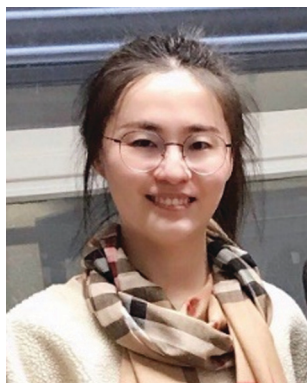
by the revolutionary Haber–Bosch process, in which pure H₂ and N₂ are reacted under high pressure at high temperature over Fe-based catalysts, which consumes 1–2% of total global fossil fuels and releases around 300 million tons of planet-warming CO₂ into the atmosphere annually.^{4–7} The development of environmentally friendly, sustainable strategies with high efficiency of NH₃ production under mild conditions is highly desirable but challenging.

Currently, there are a few green routes reported for ammonia synthesis. One is biological nitrogen fixation which mostly relies on diazotrophs in nature.⁸ Another one is electrochemical reduction of N₂, which uses electricity to produce NH₃.^{9,10} Compared with the stringent conditions of biological nitrogen fixation and the requirement of highly conductive electrolytes

^a Key Laboratory of Cluster Science, Ministry of Education of China, Key Laboratory of Photoelectronic/Electrophotonic Conversion Materials, School of Chemistry and Chemical Engineering, Beijing Institute of Technology, Beijing 100081, China

^b Department of Chemical Engineering, University College London, London WC1E 7JE, UK. E-mail: junwang.tang@ucl.ac.uk

[†] Co-first authors: These authors contributed equally to this work.



Qing Han

Qing Han is an Assistant Professor of Chemistry at Beijing Institute of Technology (Beijing, China). She received her PhD in chemistry from Beijing Institute of Technology. Her research focuses on (i) the design and synthesis of nanocatalysts for photocatalysis and electrocatalysis, and (ii) device fabrication of carbon-based nanomaterials for energy conversion.



Haimiao Jiao

Haimiao Jiao received his BEng and MSc in Chemical Engineering from Tianjin University and Imperial College London, respectively. Currently, he is a PhD student in Prof. Junwang Tang's group at the Department of Chemical Engineering, University College London. His research interests focus on photocatalytic hydrogen production from methanol and nitrogen reduction.



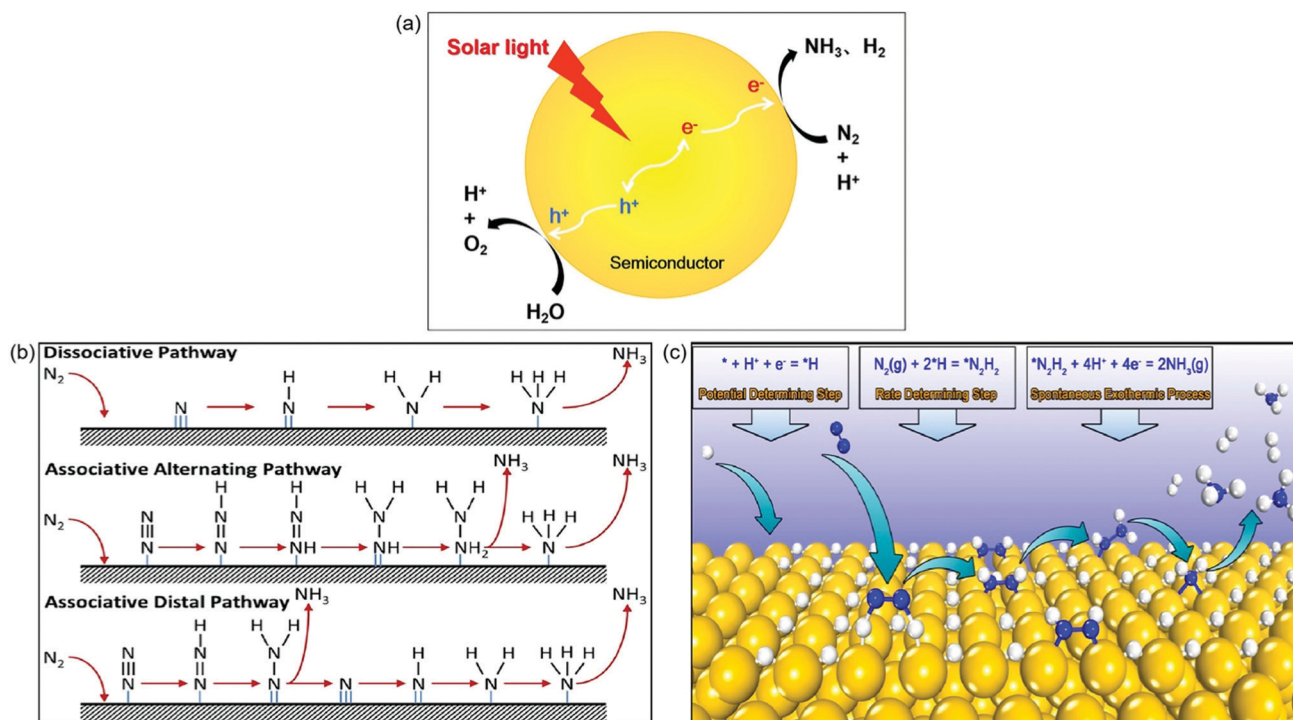


Fig. 1 (a) Photocatalytic synthesis of NH_3 in water. (b) The proposed mechanisms for photocatalytic N_2 reduction consisting of the dissociative pathway, the associative alternating pathway, and the associative distal pathway. Reproduced from ref. 27 with permission from ACS Publication. (c) The proposed mechanism of surface-hydrogenation over noble metal doped photocatalysts for NRR. Reproduced from ref. 28 with permission from ACS Publication.

Up to now, the mechanisms for N_2 reduction to NH_3 on photocatalysts have been divided into two categories based on the surface adsorption characteristics of catalysts: *N*-hydrogenation (Fig. 1b) and surface-hydrogenation (Fig. 1c).^{27,28} For the *N*-hydrogenation mechanism, there are two different pathways for NRR, dissociative and associative pathways. In the dissociative pathway, the N_2 triple bond is broken before hydrogenation, followed by hydrogenation of the N-atoms into NH_3 on the catalyst surface. The Haber-Bosch process is proved to be the dissociative pathway, which requires prohibitively high energy to cleave the $\text{N}\equiv\text{N}$ bond. Compared with the dissociative pathway, the associative pathway without the breaking of the $\text{N}\equiv\text{N}$ bond is generally used and accepted far more in the photocatalytic NRR. The nitrogen surface adsorption configuration determines the associative pathways for NRR; as shown in Fig. 1b, the hydrogenation of the adsorbed N_2 preferentially occurs at the terminal-site N owing to its relatively low steric hindrance for both the associative alternative pathway and the associative distal pathway. Having generated the first NH_3 , the other N begins a new hydrogenation circulation to produce a second NH_3 . It is actually impossible to identify the N_2 configurations due to the fact that they are favored equally by the catalysts.

Recently, a novel surface-hydrogenation mechanism for NRR has been proposed (Fig. 1c),²⁸ in which the adsorbed H^+ is reduced into $^*\text{H}$ first, and then N_2 molecules react with surface $^*\text{H}$ to generate $^*\text{N}_2\text{H}_2$, and finally the formed $^*\text{N}_2\text{H}_2$

intermediates are reduced into NH_3 . The first step is the trigger step, which enables the NRR to occur at a low potential. Furthermore, the second step is the rate determining step because it needs to conquer a high energy barrier (E_a) to break the $\text{N}\equiv\text{N}$ bond. The E_a of hydrogen evolution reaction is always lower than that of NRR, and therefore the NH_3 production rate is relatively low.

As a complete cycle indicated in Fig. 1a is the ideal process but very challenging, in most cases (or in the majority of the reported cases), a half reaction was investigated in which N_2 was reduced to NH_3 in the presence of an efficient hole scavenger which was oxidized other than water oxidation. This half reaction is relatively easy and can be used to preliminarily scan different photocatalysts for N_2 reduction. Bear in mind that this half reaction has advantages but does not indicate whether the photocatalyst selected is good for a complete cycle as indicated in Fig. 1a. The subsequent discussion will underline this in order to remind the audience of the large difference between a half reaction and the complete cycle.

Determination of NH_3

NH_3 detection is a necessary subject for the photocatalytic synthesis of NH_3 , which is of great significance in the design of high performance photocatalysts. There are four common methods of NH_3 analysis such as colorimetric assays, ^{15}N



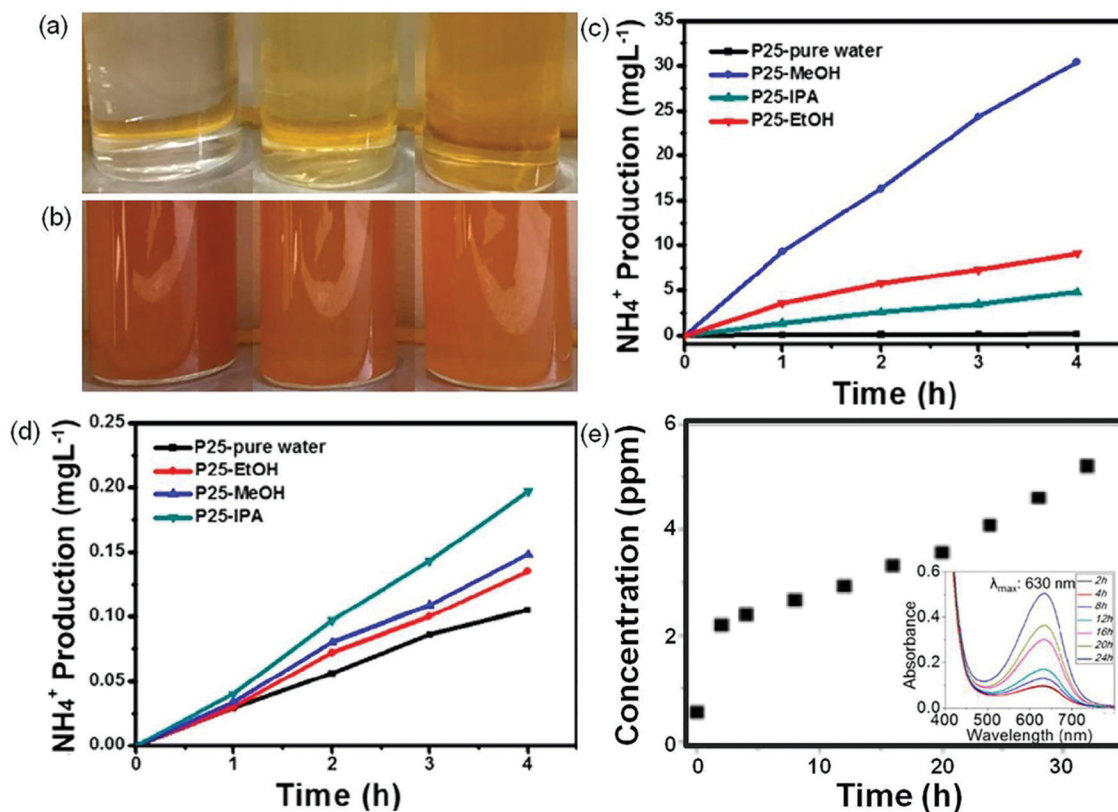


Fig. 2 Photographs of NH_3 solution with different concentrations (0, 1.2, and 4 mg L^{-1} from left to right) when (a) mixing with Nessler's reagent and (b) adding $4 \mu\text{g L}^{-1}$ formaldehyde into the above ammonia solution with Nessler's reagent. Measured results of NH_3 concentration using (c) Nessler's reagent and (d) ion chromatography methods. Reproduced from ref. 29 with permission from ACS Publication. (e) The NH_3 concentration evolution with time tested by ion chromatography. Inset in (e): UV-vis absorption spectra tested with the indophenol reagent with the characteristic adsorption peak of 630 nm wavelength. Reproduced from ref. 30 with permission from ACS Publication.

isotope labeling, ion chromatography, and *in situ* infrared spectroscopic measurements.¹² The colorimetric assays include Nessler, phenate, and indophenol blue tests for NH_3 detection. Among the colorimetric assays, Nessler's reagent has been widely used in NH_3 detection due to its advantages of simplicity and low cost.²⁹ However, it is reported that the sacrificial agents (*i.e.*, alcohol) used in the photocatalytic synthesis of NH_3 systems can be oxidized to carbonyl compounds, which will disturb NH_3 detection in Nessler's reagent detection, thus yielding misleading results. Fig. 2a shows photographs of a solution containing Nessler's reagent and different concentrations of NH_3 solutions (0, 1.2, and 4 mg L^{-1}).²⁹ It can be seen that the solution color is turning from light yellow to brown with the increase in the concentration of NH_3 . When $4 \mu\text{g L}^{-1}$ of formaldehyde is added, all the above solutions change into brown, indicating that trace formaldehyde can cause a remarkable color change (Fig. 2b). Therefore, the method of Nessler's reagent is not able to detect NH_3 production in the presence of this or similar sacrificial agents. In contrast, the generated carbonyl compounds have no effect on the testing method of ion chromatography (Fig. 2c and d). It is strongly recommended that the colorimetric assay methods have to be accompanied by other methods such as ion chromatography, ^{15}N isotope labeling, and *in situ* infrared spectroscopic measurements, to

avoid such misleading results. Fig. 2e shows the photocatalytic performance of NH_3 synthesis *via* ion chromatography and the indophenol test.³⁰ The UV-vis spectra of the reaction solution containing NH_3 exhibit a maximum absorption at 630 nm (inset in Fig. 2e), which is ascribed to the indophenols. The production of NH_3 measured by ion chromatography increases with the reaction time (Fig. 2e), consistent with the results obtained by the indophenol test, indicating that indophenol is to some extent reliable.

Extensive studies on the determination of NH_3 have revealed that the indophenol blue method, Nessler's reagent method, and the ion chromatography method are accurate when the concentration of NH_3 is less than $500 \mu\text{g L}^{-1}$, but the indophenol blue method tends to be less accurate at higher concentrations or under acidic conditions. Besides, the presence of carbonyl compounds from scavengers will enhance light adsorption in Nessler's reagent method, thus leading to an interfered detection of NH_3 . Thus, to guarantee accuracy and reliability of NH_3 quantification, at least two detection methods such as the ion chromatography method and Nessler's reagent method are recommended. Furthermore, because of trace NH_3 in the ambient atmosphere, it is necessary to test the isotopically labeled $^{15}\text{N}_2$ to confirm the source of nitrogen in the product.



Strategies to improve photocatalytic efficiency

As mentioned above there are a few reviews summarizing the activity of diverse photocatalysts.^{6,13,18,19,22} This critical review will not overlap with these reviews, but will only concentrate on the efficient strategies to improve (i) N₂ adsorption and (ii) charge separation. Such strategies include surface engineering and interface engineering for NH₃ synthesis. The photocatalytic process of NH₃ production generally includes four steps, photo-generation of charge carriers followed by their migration to the active sites, N₂ chemisorption, surface photo-oxidation and photo-reduction, and the desorption NH₃ from photocatalysts. Following the above key steps, photocatalysts can be engineered to greatly promote their catalytic activities toward NH₃ synthesis. As steps (ii), (iii) and (iv) all take place on the surface, through the study of surface regulations and interface modulations on the photocatalysts, their photocatalytic performance can be maneuvered. The surface regulations, including defect engineering and morphology engineering, significantly determine the photocatalytic activity because they can promote the surface adsorption and activation of N₂.^{23,31–40,54–57,64–71} The interface modulations, including modification with cocatalysts and semiconductors, greatly affect the charge transfer and separation efficiency.^{89–101,109–111} We will discuss these factors one by one below. In order to provide an intuitively systematic comparison of photocatalytic N₂ reduction systems, we list photocatalysts, reactants, NH₃ production rate, in particular the apparent quantum efficiency and detection method as shown in Table 1.

Surface regulations

Defect engineering. It is widely reported that vacancies in a photocatalyst can improve N₂ adsorption, to some extent even charge separation. There are different populations of vacancies, including oxygen vacancies, nitrogen vacancies, carbon vacancies, as well as sulfur vacancies.

Oxygen vacancies. The introduction of oxygen vacancies (OVs) can effectively facilitate N₂ adsorption and activation, as well as induce the generation of mid-gap states to promote the separation of photogenerated charge carriers.^{31–42} For instance, Zhang *et al.* reported that BiOBr nanosheets with OVs on the exposed {001} facets (BOB-001-OV) could stretch the N≡N bonds from 1.078 Å to 1.133 Å, which gave strong evidence for N₂ activation over the OVs.³¹ Theoretical calculations showed that the OV induced defect states could serve as the electron acceptor (Fig. 3a) to effectively suppress the recombination of electrons/holes and enhance the charge transfer from BOB-001-OV to the N₂ molecule. Fluorescence spectroscopy showed that the average lifetime (τ) of BOB-001-OV was 2.15 ns, around two times higher than that of BiOBr without OVs (BOB-001-H, Fig. 3b), proving that the OVs could promote the migration of charge carriers. With the function of OVs, BOB-001-OV showed a NH₃ yield rate of 104.3 $\mu\text{mol g}^{-1} \text{h}^{-1}$ under visible light illumination in the absence of organic scavengers

and noble-metal cocatalysts (Fig. 3c), and an apparent quantum efficiency (AQE) of 0.23% at 420 nm. Meanwhile, the amount of generated O₂ was stoichiometrically approaching 3/4 of the produced NH₃, which proved that water could act as an electron donor to achieve a complete chemical cycle in NH₃ photosynthesis.

Apart from the research on bismuth oxyhalide materials, Shiraishi's group utilized a series of commercial TiO₂ to study the effect of OVs on the photoreduction of N₂ to NH₃.³² A commercial TiO₂ of JRC-TIO-6 showed an excellent performance with a steadily increasing NH₃ production within 100 h under UV light irradiation in pure water. They proposed that the Ti³⁺ species formed at the OVs on the surface of TiO₂ served as active sites, which could effectively trap the electrons and promote N₂ dissociation, resulting in a photocatalytic NH₃ production rate of 2.5 $\mu\text{mol h}^{-1}$.

In order to create the OVs, most studies used oxygen-containing catalysts. In recent years, layered double hydroxide (LDH, Fig. 4a) materials have attracted substantial attention due to their tunable electron structure and low cost.³⁴ Zhang's group firstly used the CuCr-LDH nanosheets (CuCr-NS) with OVs as efficient photocatalysts for NH₃ photosynthesis.³⁴ The OVs over CuCr-NS surface were created by the reduction of thickness and the in-plane compressive strain (Fig. 4b). The optimal CuCr-NS exhibited a NH₃ production rate of 78.6 $\mu\text{mol g}^{-1} \text{h}^{-1}$ under full-spectrum irradiation (Fig. 4c), and an AQE of 2.4% at 400 nm. Subsequently, they further reported another work on ZnAl-LDH nanosheets for optimal OVs by incorporating coordinately unsaturated Cu^{δ+} species.³⁵ The DFT calculations implied that the OVs and Cu^{δ+} species in ZnAl-LDH could effectively promote N₂ adsorption and activation, leading to a remarkable NH₃ yield of 110 $\mu\text{mol L}^{-1} \text{h}^{-1}$ in pure water under UV-vis irradiation. Therefore, the creation of OVs on the surface of photocatalysts paves a way for developing other innovative materials toward photocatalytic N₂ reduction.

Nitrogen vacancies/carbon vacancies. Inspired by the conclusion of N₂ activation and efficient electron transfer by OVs on the catalyst surface, nitrogen vacancies (NVs) were later studied for the NRR process. Because of its excellent optical properties, low cost, and good stability, graphitic carbon nitride (g-C₃N₄) has attracted much scientific interest in the field of photocatalysis.^{42–53} Dong *et al.* have reported the application of NV-incorporated g-C₃N₄ (V-g-C₃N₄) for N₂ photoreduction.⁵⁴ They found that the N₂ adsorption amount of V-g-C₃N₄ was 75.1 $\text{cm}^3 \text{g}^{-1}$, 2.4 times higher than that of g-C₃N₄ without NVs (30.9 $\text{cm}^3 \text{g}^{-1}$), which suggested that the NVs on the surface of V-g-C₃N₄ could provide a number of chemisorption and activation sites for N₂ molecules, leading to an NH₃ production rate of 1.24 $\text{mmol h}^{-1} \text{g}_{\text{cat}}^{-1}$ under visible light irradiation in the presence of methanol as a hole scavenger. Recently, a g-C₃N₄ catalyst modified with cyano groups (–C≡N) and intercalated K⁺ (mCNCN) was prepared as an efficient photocatalyst for NRR.⁵⁵ Relative to pristine g-C₃N₄, the absorption spectrum of mCNCN was extended to the whole visible region (Fig. 5a), resulting from the π – π^* electronic transitions in the conjugated heterocyclic rings and the n – π^* electronic transitions at the



Table 1 Photocatalytic performance of different catalysts for NH₃ synthesis

Catalyst	Reaction medium	Scavenger	Light source	Ammonia yield	Apparent quantum efficiency (AQE)	Ammonia detection method	Ref.
BiOBr-001-OV	N ₂ /H ₂ O (l)	None	$\lambda > 420$ nm	104.3 $\mu\text{mol g}^{-1} \text{h}^{-1}$	0.23% at 420 nm	Nessler's reagent	31
JRC-TIO-6	N ₂ /H ₂ O (l)	2-PrOH	$\lambda > 280$ nm	2.5 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	32
CuCr-LDH	N ₂ /H ₂ O (l)	None	Full spectrum	78.6 $\mu\text{mol g}^{-1} \text{h}^{-1}$	2.4% at 400 nm	Nessler's reagent	34
ZnAl-LDH	N ₂ /H ₂ O (l)	None	UV-vis	110 $\mu\text{mol L}^{-1} \text{h}^{-1}$	1.77% at 265 nm, 0.56% at 365 nm	Ion chromatography	35
V-g-C ₃ N ₄	N ₂ /H ₂ O (l)	Methanol	$\lambda > 420$ nm	1.24 $\text{mmol h}^{-1} \text{g}_{\text{cat}}^{-1}$	Not reported	Nessler's reagent	54
mCNN	N ₂ /ethylene glycol (l)	Ethylene glycol	$\lambda > 400$ nm	3.42 $\text{mmol g}^{-1} \text{h}^{-1}$	Not reported	Nessler's reagent	55
SCNNs	N ₂ /H ₂ O (l)	Methanol	Full spectrum	5.99 $\text{mmol h}^{-1} \text{g}_{\text{cat}}^{-1}$	Not reported	Nessler's reagent	56
Mo _{0.1} Ni _{0.1} Cd _{0.8} S	N ₂ /H ₂ O (l)	Ethanol	UV	3.2 $\text{mg h}^{-1} \text{g}_{\text{cat}}^{-1}$	Not reported	Nessler's reagent	57
0.2 wt% Fe-doped TiO ₂	N ₂ /H ₂ O (g)	None	UV	11.6 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	17
0.4 wt% Co-doped TiO ₂	N ₂ /H ₂ O (l)	None	UV	6.3 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	17
0.4 wt% Cr-doped TiO ₂	N ₂ /H ₂ O (l)	None	UV	0.37 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	17
0.4 wt% Mo-doped TiO ₂	N ₂ /H ₂ O (l)	None	UV	6.7 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	17
Fe-Doped TiO ₂	N ₂ /H ₂ O (l)	Ethanol	$\lambda = 254$ nm	1.2 mmol L^{-1}	Not reported	Nessler's reagent	64
Fe-Doped SrMoO ₄	N ₂ /H ₂ O (l)	None	UV-vis	93.1 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Nessler's reagent, indophenol blue method	65
Fe-Doped BiOCl	N ₂ /H ₂ O (l)	None	Full spectrum	1.022 $\text{mmol g}^{-1} \text{h}^{-1}$	1.8% at 420 nm	Indophenol blue method	66
Mo-Doped W ₁₈ O ₄₉	N ₂ /H ₂ O (l)	Na ₂ SO ₃	Full spectrum	195.5 $\mu\text{mol g}^{-1} \text{h}^{-1}$	0.33% at 400 nm	Nessler's reagent, ion chromatography	67
Cu-Doped TiO ₂	N ₂ /H ₂ O (l)	None	Full spectrum	78.9 $\mu\text{mol g}^{-1} \text{h}^{-1}$	0.08% at 600 nm, 0.05% at 700 nm	Nessler's reagent, ion chromatography	68
FePt@C ₃ N ₄	N ₂ /H ₂	None	$\lambda > 400$ nm	63 $\mu\text{g h}^{-1} \text{g}^{-1}$	0.15% between 450 and 500 nm	Colorimetric method, ion chromatography	72
B-Doped g-C ₃ N ₄	N ₂ /H ₂ O (l)	Na ₂ SO ₃	$\lambda > 400$ nm	313.9 $\mu\text{mol g}^{-1} \text{h}^{-1}$	0.64% at 420 nm	Nessler's reagent	73
C-Doped TiO ₂	N ₂ /H ₂ O (l)	Methanol	$\lambda > 420$ nm	109.3 $\mu\text{mol g}^{-1} \text{h}^{-1}$	2.4% at 400 nm	Ion chromatography	77
Bi ₅ O ₇ I-001	N ₂ /H ₂ O (l)	Methanol	280–800 nm	111.5 $\mu\text{mol g}^{-1} \text{h}^{-1}$	5.1% at 365 nm	Nessler's reagent	77
BiOCl-010	N ₂ /H ₂ O (l)	Methanol	Full spectrum	92.4 $\mu\text{mol g}^{-1} \text{h}^{-1}$	4.3% at 254 nm	Nessler's reagent	78
BiO	N ₂ /H ₂ O (l)	None	Full spectrum	1226 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	86
Bi ₅ O ₇ Br-NT	N ₂ /H ₂ O (l)	None	$\lambda > 400$ nm	1.38 $\text{mmol g}^{-1} \text{h}^{-1}$	2.3% at 420 nm	Nessler's reagent	87
eBP NFs	N ₂ /H ₂ O (l)	Na ₂ SO ₃ , Na ₂ S·9H ₂ O	$\lambda = 420$ nm	2.37 $\text{mmol g}^{-1} \text{h}^{-1}$	Not reported	Nessler's reagent	88
SiO ₂ /C-RP	N ₂ /H ₂ O (l)	None	Full spectrum	0.73 $\mu\text{mol h}^{-1}$	Not reported	Ion chromatography	89
Ru-Loaded TiO ₂	N ₂ /H ₂ O (l)	Methanol	Full spectrum	29.4 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	93
Ru-Loaded TiO ₂	N ₂ /H ₂ O (l)	Ethanol	Full spectrum	56.3 $\mu\text{g h}^{-1} \text{g}_{\text{cat}}^{-1}$	Not reported	Indophenol blue method	95
Au/TiO ₂ -OV	N ₂ /H ₂ O (l)	Methanol	$\lambda > 420$ nm	78.6 $\mu\text{mol g}^{-1} \text{h}^{-1}$	0.82% at 550 nm	Indophenol blue method	96
Ti ₃ C ₂ -P25	N ₂ /H ₂ O (l)	None	Full spectrum	10.74 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Ion chromatography	103
Pt-Loaded ZnO	N ₂ /H ₂ O (l)	Na ₂ SO ₃	UV	860 $\text{mmol g}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	104
SV-1T-MoS ₂ /CdS	N ₂ /H ₂ O (l)	Methanol	420–780 nm	8220.83 $\mu\text{mol L}^{-1} \text{h}^{-1} \text{g}^{-1}$	4.424% under simulated solar light	Nessler's reagent	105
NiS/CdS	N ₂ /H ₂ O (l)	None	Full spectrum	2.8 $\text{mg L}^{-1} \text{h}^{-1}$	0.76% at 420 nm	Nessler's reagent	106
Pt-SACs/CTF	N ₂ /H ₂ O (l)	None	420–780 nm	171.4 $\mu\text{mol g}^{-1} \text{h}^{-1}$	1.4% at 420 nm	Nessler's reagent	107
Ru/RuO ₂ /g-C ₃ N ₄	N ₂ /H ₂ O (l)	Methanol	Full spectrum	13.3 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Ion chromatography	108
Fe@graphene	N ₂ /H ₂ , 200 °C	None	200–600 nm	420 $\mu\text{g g}_{\text{cat}}^{-1} \text{h}^{-1}$	Not reported	Indophenol blue method	110
g-C ₃ N ₄ /Bi ₂ MoO ₆	N ₂ /H ₂ O (l)	Methanol	$\lambda > 420$ nm	3271 $\mu\text{mol L}^{-1} \text{g}^{-1}$	Not reported	Nessler's reagent	111
TiO ₂ @C/g-C ₃ N ₄	N ₂ /H ₂ O (l)	Methanol	$\lambda > 420$ nm	250.6 $\mu\text{mol g}^{-1} \text{h}^{-1}$	0.14% at 420 nm	Nessler's reagent	112
Bi ₂ MoO ₆ /OV-BiOBr	N ₂ /H ₂ O (l)	None	Full spectrum	81.0 $\mu\text{mol g}^{-1} \text{h}^{-1}$	Not reported	Nessler's reagent	116

active N defect sites ($-\text{C}\equiv\text{N}$), which indicated that the introduction of the N-defect of $-\text{C}\equiv\text{N}$ could significantly change

the optical properties. Fig. 5b displays the band structures of mCNN and pristine g-C₃N₄ through the results of experiments.



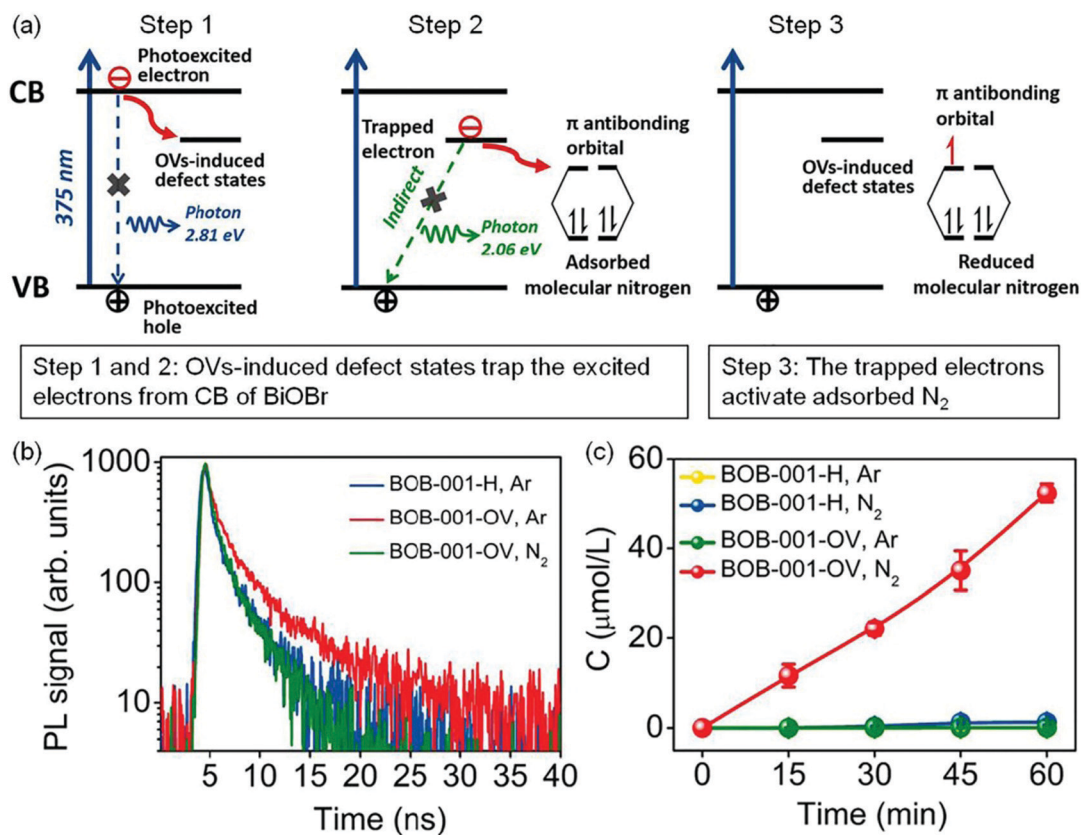


Fig. 3 Schematic of the OV-induced enhanced interfacial electron transfer processes over BOB-001-OV. (b) Time-resolved photoluminescence (PL) spectra of BOB-001-OV and BOB-001-H. (c) The photocatalytic NH_3 production of BOB-001-OV and BOB-001-H under visible light ($\lambda > 420$ nm). Reproduced from ref. 31 with permission from ACS Publication.

Compared with pristine $g\text{-C}_3\text{N}_4$, a subgap state was presented for mCNN, which could be ascribed to the N defect, resulting in a broad visible-light response. The enhanced EPR intensity (Fig. 5c) of mCNN compared to pristine $g\text{-C}_3\text{N}_4$ was assigned to the strong electron withdrawing groups of $-\text{C}\equiv\text{N}$ on mCNN, which could delocalize the isolated electrons in mCNN, therefore promoting the separation of photocarriers and increasing the generation of active radical species to boost the activity of NH_3 photosynthesis. The rate of NH_3 generation of the mCNN photocatalyst was measured to be $3.42 \text{ mmol g}^{-1} \text{ h}^{-1}$ with ethylene glycol as a scavenger under visible light irradiation (Fig. 5d), which was much higher than that of pristine $g\text{-C}_3\text{N}_4$ ($1.11 \text{ mmol g}^{-1} \text{ h}^{-1}$). The DFT calculations revealed that K^+ could be linked to the unsaturated C centers by coordination as shown in Fig. 5e, which could adsorb N_2 as the lone-pair electrons of nitrogen would fill the empty state in K^+ ($\Delta G = -0.28 \text{ eV}$), followed by the rearrangement of K^+ with the C atom and the adsorption of N_2 ($\Delta G = -0.17 \text{ eV}$). Because of the large ion hydration free energy of K^+ , it could be extracted back into solution, then forming a C_2N_4 ring ($\Delta G = 1.64 \text{ eV}$) intermediate which could undergo a MvK process to regenerate $-\text{C}\equiv\text{N}$, thereby stabilizing the unsaturated C sites.

Apart from the above NVs on the $g\text{-C}_3\text{N}_4$ photocatalyst, carbon vacancies (CVs) on a porous sulfur-doped $g\text{-C}_3\text{N}_4$ (SCNNSs) have also been exploited.⁵⁶ The introduction of CVs

was proved to play the same role as NVs, while the S-doping effect was demonstrated to improve the separation efficiency of photogenerated electron-hole pairs. Taking these advantages, the optimal SCNNSs showed an NH_3 production rate of $5.99 \text{ mmol h}^{-1} \text{ g}_{\text{cat}}^{-1}$.

Sulfur vacancies. Considering the importance of the sulfur element in nitrogenases, a few have reported the effects of sulfur vacancies (SVs) on N_2 photoreduction. It was reported that SVs could be created on the surface of $\text{Mo}_{0.1}\text{Ni}_{0.1}\text{Cd}_{0.8}\text{S}$ by co-doping Mo and Ni into CdS.⁵⁷ N_2 -temperature-programmed desorption (N_2 -TPD) results illustrated that the SVs on the surface of $\text{Mo}_{0.1}\text{Ni}_{0.1}\text{Cd}_{0.8}\text{S}$ obviously enhanced N_2 chemisorption (Fig. 6a), which was beneficial for NRR. Importantly, the photocatalytic NH_3 generation rate of $\text{Mo}_{0.1}\text{Ni}_{0.1}\text{Cd}_{0.8}\text{S}$ was linearly related to its SV concentration as shown in Fig. 6b. As the concentration of SVs increased, there was an increase in the photocatalytic NH_3 production rate. This result suggested that the SVs played a vital role in the photoreduction of N_2 . The $\text{Mo}_{0.1}\text{Ni}_{0.1}\text{Cd}_{0.8}\text{S}$ with the highest SV concentration showed the highest NH_3 production rate of $3.2 \text{ mg h}^{-1} \text{ g}_{\text{cat}}^{-1}$ (Fig. 6b).

Doping. Metal doping is a very promising strategy to change the electronic structure and surface property of photocatalysts to enhance the photocatalytic activity.^{58–63} In one report,



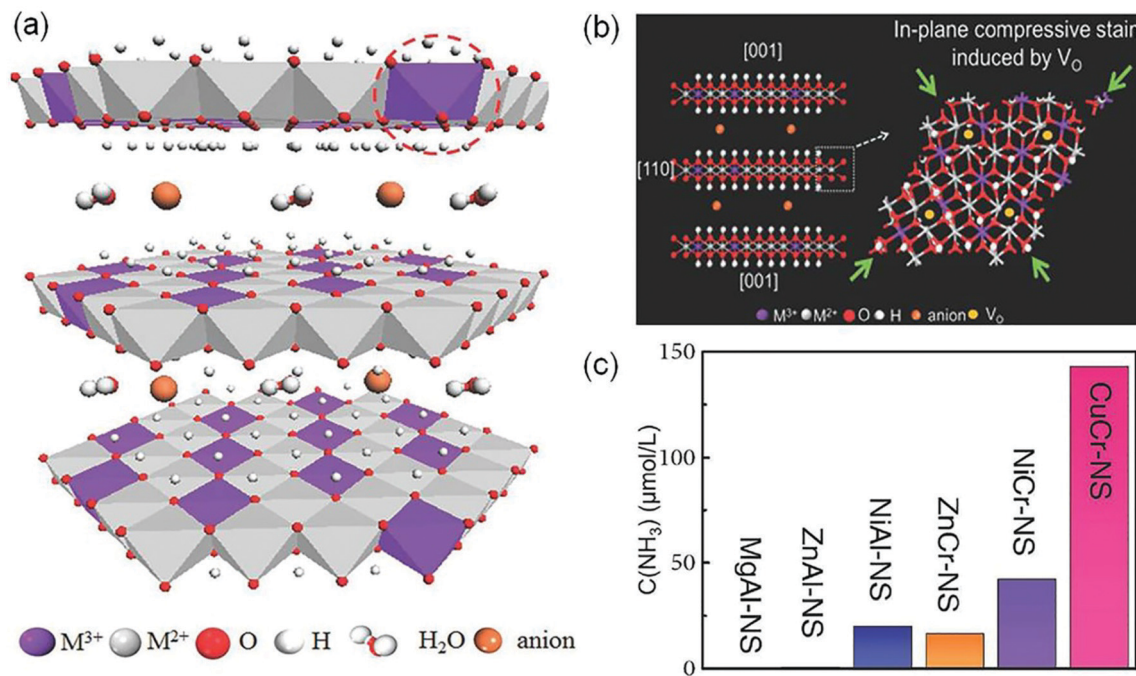


Fig. 4 (a) Schematic representation of the layered LDH structure with defective MO_6 octahedra. (b) Schematic of the in-plane biaxial compressive strain in CuCr-NS. (c) The yield of NH_3 over different LDH photocatalysts under visible-light illumination ($\lambda > 400 \text{ nm}$). Reproduced from ref. 34 with permission from Wiley-VCH.

Schrauzer *et al.* studied the effect of different metal doped TiO_2 on the photocatalytic activity of N_2 reduction, including iron (Fe), chromium (Cr), cobalt (Co), and molybdenum (Mo).¹⁷ They found that Fe doping was the most effective for enhancing the photocatalytic activity of TiO_2 for photosynthesis of NH_3 from N_2 and H_2O . However, O_2 production as the oxidation product was not reported. In this regard, Zhao *et al.* investigated the Fe-doped TiO_2 with highly exposed (101) facets for N_2 photofixation,⁶⁴ and found that an optimum Fe^{3+} doping content played a key role in inhibiting the recombination of photoinduced electron-hole pairs, which could act as a temporary electron/hole trapping sites, therefore enhancing the concentration of charge carriers and improving the photocatalytic performance. With this in mind, Liu *et al.* also reported a Fe-doped $SrMoO_4$ (FSMO) as a potential candidate for N_2 photoreduction.⁶⁵ Further studies revealed that the intrinsic bandgap of $SrMoO_4$ could be shrunk from 3.98 eV to 2.93 eV with the increase in Fe doping concentration (from 0 to 5.1%), resulting in the extension of light adsorption from the ultraviolet to the visible-light region. Besides that, the Fe doping could induce the formation of surface defects as active sites for N_2 adsorption and significantly retard the recombination of electrons and holes, leading to enhanced N_2 reduction reaction. As a result of these properties, an improved NH_3 production rate of $93.1 \mu\text{mol g}^{-1} \text{ h}^{-1}$ over the optimal FSMO was achieved compared with that of pristine $SrMoO_4$ ($66.7 \mu\text{mol g}^{-1} \text{ h}^{-1}$). In another related work incorporating Fe into BiOCl nanosheets, Fe-doped BiOCl nanosheets (BiOCl NSs-Fe) were developed for N_2 photoreduction.⁶⁶ The optimal BiOCl NSs-Fe exhibited a marked enhancement of photocatalytic NH_3 production, and

the efficiency was 2.53 times higher than that of pristine BiOCl NSs.

In another study, Mo was successfully doped into $W_{18}O_{49}$ nanowires to produce Mo-doped $W_{18}O_{49}$ nanowires (MWO-1) as shown in Fig. 7a.⁶⁷ Compared with OV-rich $W_{18}O_{49}$ nanowires, Mo-doping had many kinds of effects on photocatalytic N_2 reduction. Fig. 7b displays the electronic band structures of MWO-1 and $W_{18}O_{49}$; it is observed that the defect-band center could be shifted to the Fermi level by Mo doping, which provided more energetic electrons for NRR. Theoretical simulations revealed that the Mo-W centers could alter electron distribution, leading to a larger adsorption energy (-2.48 eV) relative to $W_{18}O_{49}$ with W-W sites (1.65 eV), which was in favor of N_2 chemisorption and activation (Fig. 7c). Additionally, the enhanced M-O co-valence caused by Mo doping could effectively promote the electron transfer from metal active centers to adsorbed N_2 molecules. Thus, the optimal MWO-1 exhibited an excellent NH_3 yield rate of $195.5 \mu\text{mol g}^{-1} \text{ h}^{-1}$ under full-spectrum irradiation (Fig. 7d), which was 7 times higher than that of $W_{18}O_{49}$, and also realized a high AQE of 0.33% at 400 nm.

Furthermore, Cu-doped ultrathin TiO_2 nanosheets (TiO_2-V_O -strain) were also studied.⁶⁸ Compared with pristine TiO_2 , introducing Cu into TiO_2 could create abundant OVs and cause lattice distortion and strain effects, resulting in a significant increase of N_2 adsorption energy on the surface of TiO_2-V_O -strain (-0.37 eV) compared with that of pristine TiO_2 (-0.17 eV , Fig. 8a). The increased adsorption energy promoted the electron transfer from TiO_2 to N_2 , thereby breaking the $N \equiv N$ bonds. Meanwhile, the required reaction energy for N_2 hydrogenation to



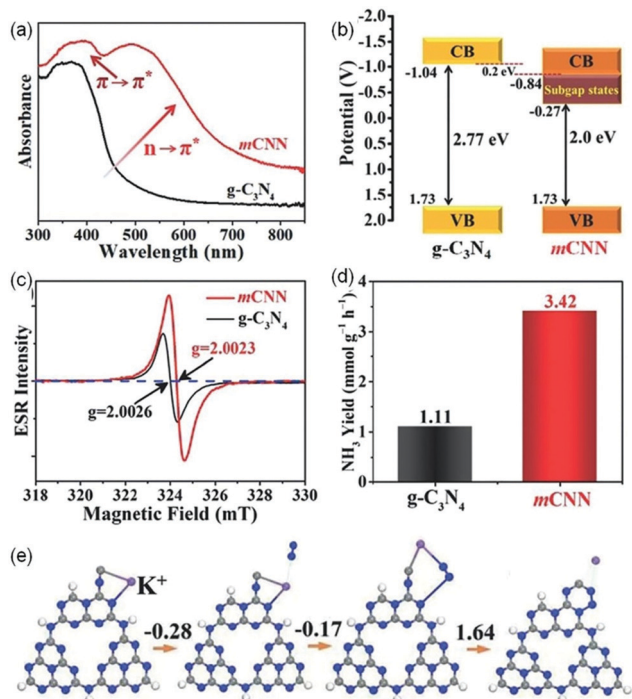


Fig. 5 (a) UV-vis diffuse reflectance spectra (DRS). (b) Schematic illustration of energy band structures. (c) Room-temperature electron paramagnetic resonance (EPR) spectra. (d) NH_3 production rates of $\text{g-C}_3\text{N}_4$ and mCNN . (e) The calculated free-energy changes for the NRR pathway using mCNN . The blue sphere represents N, the grey sphere represents C, the white sphere represents H, and the purple sphere represents K^+ . Reproduced from ref. 55 with permission from Wiley-VCH.

N-NH^* was only 0.365 eV on $\text{TiO}_2\text{-V}_\text{o}$ -strain, which was lower than that on pristine TiO_2 (2.115 eV, Fig. 8b). As a result, the optimal $\text{TiO}_2\text{-V}_\text{o}$ -strain (6% TiO_2) achieved an enhanced NH_3 yield rate of $78.9 \mu\text{mol g}^{-1} \text{h}^{-1}$ under full solar irradiation, about 5.2 times higher than that of the TiO_2 nanosheets with OV (0% TiO_2 , Fig. 8c). In addition, the yield rate of O_2 ($59.1 \mu\text{mol g}^{-1} \text{h}^{-1}$) and NH_3 mentioned above over $\text{TiO}_2\text{-V}_\text{o}$ -strain was close to the

stoichiometric ratio of 3:4, revealing that the protons in NH_3 were from the H_2O molecules. Most recently, a novel bimetallic system with Fe–Pt loaded $\text{g-C}_3\text{N}_4$ was used for efficient ammonia synthesis under mild conditions.⁷² Further investigations proposed that the doping of Pt onto the Fe nanocluster over the surface of $\text{g-C}_3\text{N}_4$ could cause an uplift of the energy band of semiconductors and form a large Schottky barrier, thus leading to an improved separation of photogenerated carriers and enhanced N_2 reduction. Accordingly, 0.3 wt% Pt doped on 3 wt% $\text{Fe@C}_3\text{N}_4$ exhibited a high NH_3 production rate of $63 \mu\text{mol g}^{-1} \text{h}^{-1}$ with gaseous H_2 and N_2 as reactants under visible light irradiation and the AQE was tested to be 0.15% between 450 and 500 nm.

Apart from metal doping, a metal-free B-doped $\text{g-C}_3\text{N}_4$ nanosheet (BCN) with exposed active N atoms was recently synthesized for highly efficient ammonia synthesis.⁷³ Theoretical studies revealed that the exposed N atoms could be stabilized by B–N–C coordination on BCN, which was different from that in the pristine $\text{g-C}_3\text{N}_4$. The B dopants were proved as active sites for N_2 adsorption and activation and could effectively retard charge recombination and improve light utilization. The optimal BCN with 13.8 wt% B-dopants exhibited a remarkable NH_3 yield rate of $313.9 \mu\text{mol g}^{-1} \text{h}^{-1}$ under visible light irradiation in the presence of Na_2SO_3 as a hole scavenger, which was much higher than that ($32.8 \mu\text{mol g}^{-1} \text{h}^{-1}$) over pristine $\text{g-C}_3\text{N}_4$ and also achieved a good QE of around 0.64% at 420 nm. In addition, carbon-doped TiO_2 was also reported for N_2 photoreduction, which exhibited an NH_3 yield rate of $109.3 \mu\text{mol g}^{-1} \text{h}^{-1}$.

Morphology engineering

Facets. Since the surface atomic distribution has a huge effect on the active sites and electronic structures of photocatalysts,^{69,74–76} many researchers attempted to control the crystal facet of semiconductors to enhance the photocatalytic N_2 reduction. For instance, two $\text{Bi}_5\text{O}_7\text{-I}$ nanosheets with different {100} and {001} facets were synthesized successfully *via* hydrolysis and calcination methods, respectively.⁷⁷ Further investigations

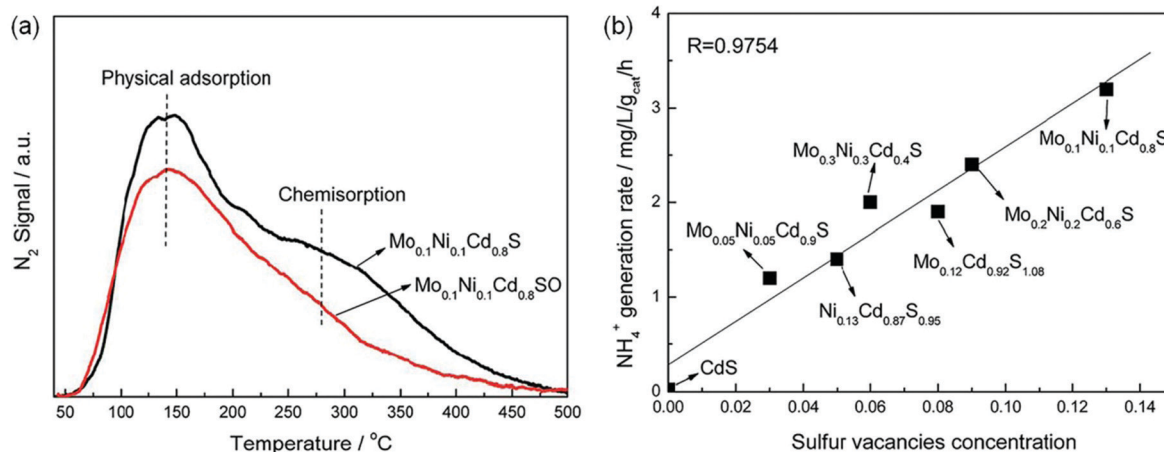


Fig. 6 (a) The N_2 -TPD of $\text{Mo}_{0.1}\text{Ni}_{0.1}\text{Cd}_{0.8}\text{S}$ and $\text{Mo}_{0.1}\text{Ni}_{0.1}\text{Cd}_{0.8}\text{SO}$. (b) The relationship of NH_3 yield over the obtained photocatalysts and the SV concentration. Reproduced from ref. 57 with permission from The Royal Society of Chemistry.



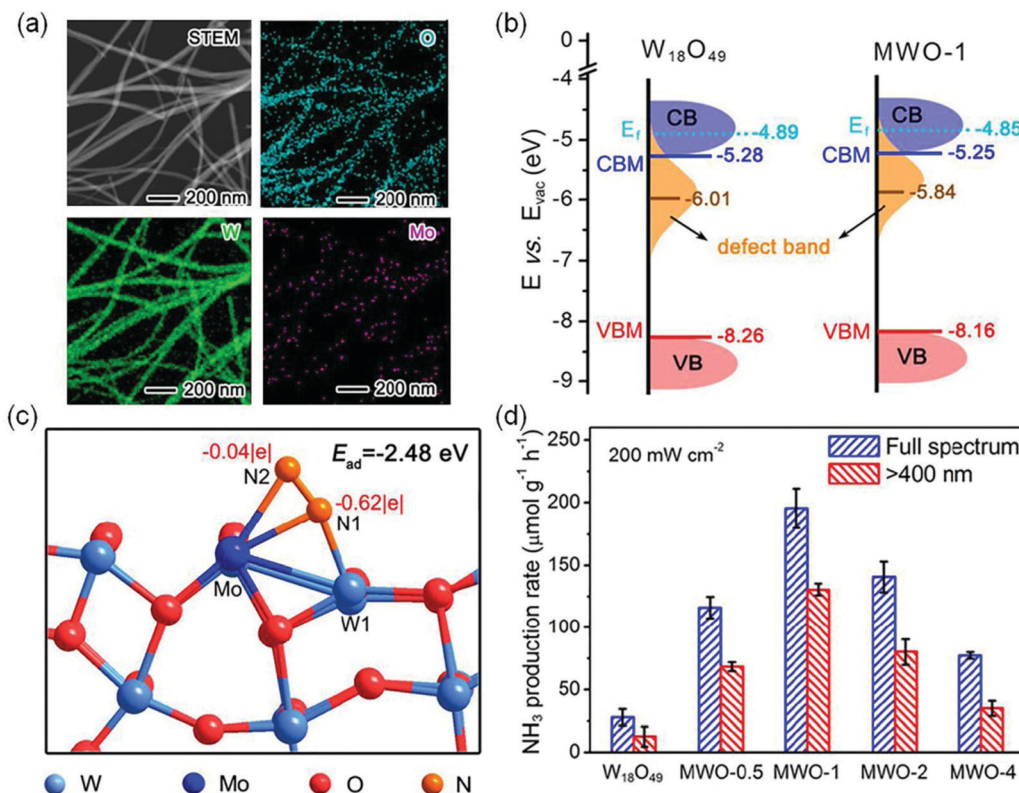


Fig. 7 (a) The scanning TEM (STEM) image and the corresponding elemental mapping of MWO-1. (b) Schematic illustration of the band structures of $W_{18}O_{49}$ and MWO-1. (c) Schematic simulation for N_2 adsorption and activation on the surface active sites of the Mo-doped $W_{18}O_{49}$ model. (d) The photocatalytic NH_3 yield on various catalysts. Reproduced from ref. 67 with permission from ACS Publication.

revealed that Bi_5O_7I-001 had more negative conduction band position (-1.45 eV) compared with that in Bi_5O_7I-100 (-0.85 eV) and exhibited higher efficiency for separation of photoinduced carriers, thus leading to a high photocatalytic activity for N_2 photoreduction. Accordingly, a remarkable NH_3 generation rate of $111.5 \mu mol g^{-1} h^{-1}$ over Bi_5O_7I-001 was achieved using methanol as a hole scavenger and the AQE of that was increased to 5.1%.

In another report, Zhang *et al.* also systematically investigated the influence of $\{001\}$ and $\{010\}$ facets of OV-rich $BiOCl$ nanosheets on N_2 adsorption and activation.⁷⁸ Experimental results showed that $\{010\}$ facet $BiOCl$ nanosheets (BOC-010) exhibited a superior performance with a NH_3 yield of $4.62 \mu mol g^{-1} h^{-1}$, which was around 2.5 times higher than that on $\{001\}$ facet $BiOCl$ nanosheets (BOC-001, Fig. 9a). Meanwhile, it was observed that N_2H_4 which served as a main intermediate was accumulated on BOC-010 during N_2 photoreduction reaction within 30 min and was consumed gradually to be converted into NH_3 (Fig. 9b). Importantly, when N_2H_4 was used as the reactant, both BOC-010 and BOC-001 showed a similar photocatalytic conversion efficiency of N_2H_4 under simulated solar irradiation (Fig. 9c), which in turn meant that BOC-010 had a stronger ability to generate N_2H_4 intermediates. Further DFT calculations revealed that the N_2 fixation on BOC-010 followed an alternative pathway with N_2H_4 as the main intermediate, which could provide a lower energy pathway for N_2

fixation *via* proton-coupled electron transfer compared with that *via* a distal pathway occurring on BOC-001 facets, thus resulting in enhanced NH_3 production (Fig. 9d and e).

Nanostructure engineering. Nanomaterials have preternatural interface structures and unique functions, such as the small size effect, the surface and boundary effect, quantum size and so on.^{79–85} In order to increase the catalytic active sites of the photocatalytic reduction of N_2 , Sun *et al.* synthesized BiO quantum dots with an average size of 2–5 nm (Fig. 10a), which showed an NH_3 production rate of $1226 \mu mol g^{-1} h^{-1}$ without a cocatalyst and a sacrificial agent.⁸⁶ Meanwhile, the amount of produced O_2 increased continuously with increasing irradiation time, consistent with the stoichiometric chemistry. Kinetic analysis and quantum chemical calculations suggested that the highly efficient photocatalytic activity of BiO quantum dots could be attributed to the fact that the smaller size endowed the surface and edge of BiO rich in Bi^{2+} species, which had the potential to increase the electron donation to the anti-bonding π^* orbitals of N_2 , thereby could act as active sites to adsorb and activate N_2 (Fig. 10b). More recently, a tubular OV-rich Bi_5O_7Br (Bi_5O_7Br-NT) with a diameter of 5 nm (Fig. 10c) was prepared and it exhibited a maximum NH_3 yield rate of $1.38 mmol g^{-1} h^{-1}$ (Fig. 10d) upon changing the OV concentration.⁸⁷ A high specific surface area of up to $96.56 m^2 g^{-1}$ was observed for Bi_5O_7Br-NT , which indicated the exposure of abundant surface



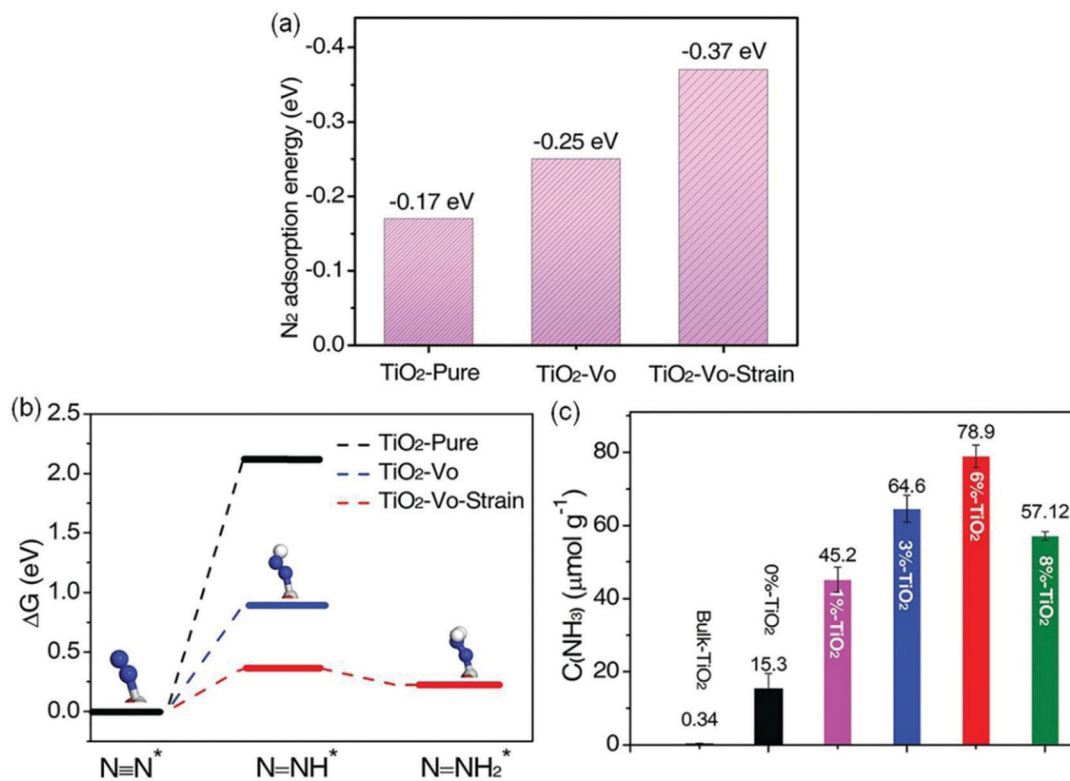


Fig. 8 (a) The N_2 adsorption energies by DFT calculations on different photocatalysts. (b) The calculated Gibbs free energy for N_2 reduction on the (001) surface of different photocatalysts (H: white sphere, Ti: gray sphere, N: blue sphere, O: red sphere). (c) Yield of NH_3 over various samples under UV-vis irradiation. Reproduced from ref. 68 with permission from Wiley-VCH.

OVs for the chemisorption and activation of N_2 . It was found that the OVs on the $\text{Bi}_5\text{O}_7\text{-Br-NT}$ surface could not only effectively promote N_2 adsorption and activation, but also could be regenerated by capturing O atoms from H_2O after the reaction to maintain the good stability of $\text{Bi}_5\text{O}_7\text{-Br}$ nanostructures. As a result, the calculated AQE for the $\text{Bi}_5\text{O}_7\text{-Br-NT}$ photocatalyst was 2.3% under visible light irradiation at 420 nm (Fig. 10e). Furthermore, it was reported that metal-free black phosphorus nanoflakes with abundant edges (eBP NFs) synthesized *via* a facile chemical etching exfoliation method exhibited a remarkable NH_3 yield rate of $2.37 \text{ mmol g}^{-1} \text{ h}^{-1}$ under visible light irradiation in the presence of Na_2SO_3 and $\text{Na}_2\text{S}\cdot 9\text{H}_2\text{O}$ as scavengers.⁸⁸ The photoelectrochemical characteristics (PEC) and transient absorption (TA) studies revealed that the efficient ammonia synthesis was attributed to rich edges on the surface, which could provide abundant active sites for enhanced N_2 adsorption and activation. Similarly, a nanocomposite ($\text{SiO}_2/\text{C-RP}$) prepared by loading red phosphorus on SiO_2 nanospheres showed superior charge separation.⁸⁹ Besides, this hybrid nanostructure exhibited a large surface area, good water dispersibility and large light adsorption. As a result of these benefits, $\text{SiO}_2/\text{C-RP}$ exhibited an NH_3 production yield of $0.73 \mu\text{mol h}^{-1}$ under full-spectrum irradiation.

Interfacial modulation

Cocatalyst loading. With the similar idea of using metal cocatalysts to inhibit the recombination of electrons and holes

and enhance the photostability of catalysts in other photocatalytic processes, e.g. water splitting,^{90–102} Ranjit *et al.* studied different metal loaded TiO_2 toward the photoreduction of N_2 , including ruthenium (Ru), rhodium (Rh), palladium (Pd), and platinum (Pt).⁹³ It was found that the order of the photocatalytic activity of these metals on TiO_2 was $\text{Ru} > \text{Rh} > \text{Pd} > \text{Pt}$. The excellent performance of Ru-loaded TiO_2 could be assigned to the high metal-hydrogen bond strength which well correlated with the high production of NH_3 . In recent years, single-atom catalysts (SACs) have been well studied as excellent candidates for boosting the catalytic activity and improving the utilization efficiency of metal atoms during the photocatalytic process.⁹⁴ To further improve the performance of the photocatalytic NH_3 synthesis, a single atomic Ru decorated TiO_2 nanosheet with abundant OVs was prepared.⁹⁵ The DFT calculations indicated that the single atomic Ru species were favorable to be formed and dispersed on the OVs over the TiO_2 nanosheet surface. Further investigations revealed that the isolated Ru atom could serve as the active site for N_2 adsorption and activation, promote the separation of photogenerated charge carriers efficiently, as well as inhibit the competitive H_2 evolution reaction, thus leading to an improved NH_3 photosynthesis. Thus, a 1 wt% single Ru atom decorated TiO_2 nanosheet exhibited a superior performance with an NH_3 yield of $56.3 \mu\text{g h}^{-1} \text{ g}_{\text{cat}}^{-1}$, which was two times as high as that of pure TiO_2 . Recently, Au-anchored TiO_2 nanosheets with OV ($\text{Au}/\text{TiO}_2\text{-OV}$) photocatalysts were developed for N_2 photofixation.⁹⁶ As shown in Fig. 11a, Au nanospheres



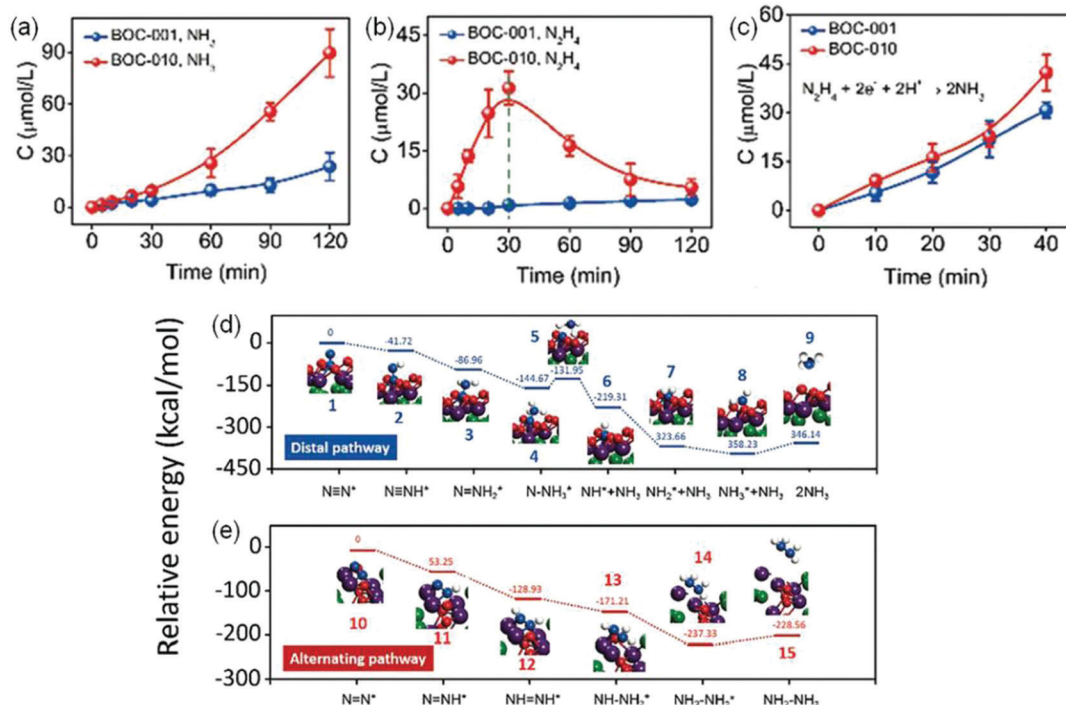


Fig. 9 Photocatalytic performance for N₂ reduction on BOC-001 and BOC-010. The yield comparison of produced (a) NH₃ and (b) N₂H₄ on BOC-001 and BOC-010 under simulated solar light irradiation. (c) The NH₃ yield produced from N₂H₄ BOC-001 and BOC-010 under simulated solar light irradiation. The calculated free energy change for N₂ reduction (d) on the (001) surface via a distal pathway and (e) on the (010) surface via an alternating pathway. Reproduced from ref. 78 with permission from The Royal Society of Chemistry.

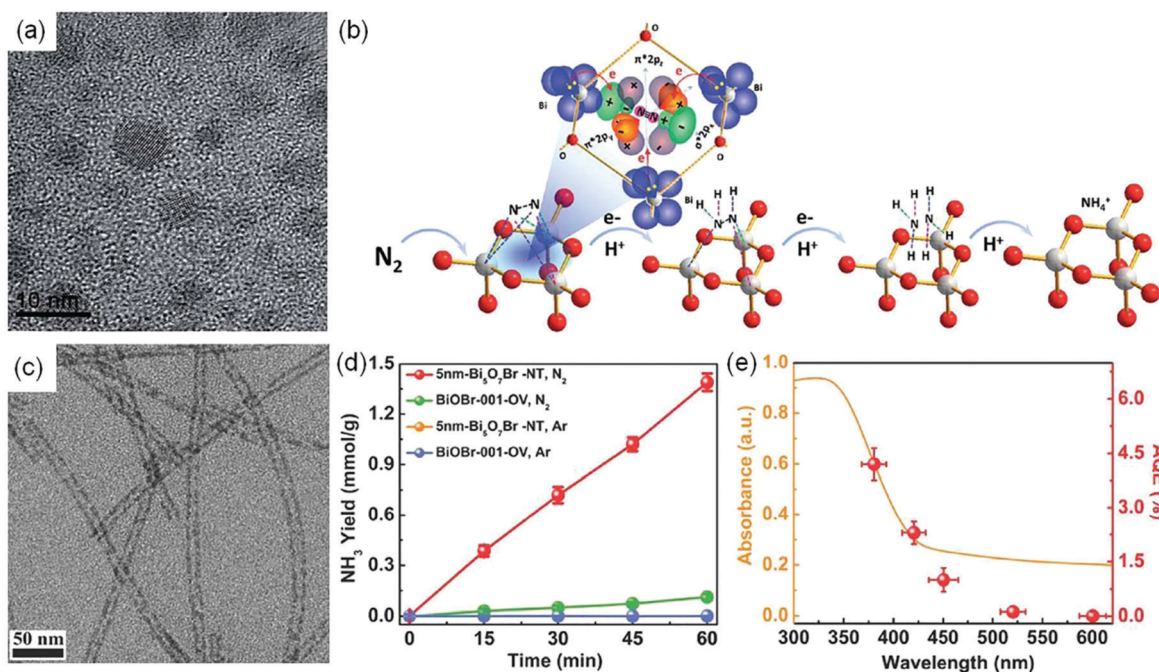


Fig. 10 (a) TEM image of BiO quantum dots. (b) Proposed mechanism for N₂ photoreduction on BiO quantum dots. Reproduced from ref. 86 with permission from The Royal Society of Chemistry. (c) TEM image of Bi₅O₇Br-NT. (d) Yield of NH₃ for Bi₅O₇Br-NT. (e) The wavelength-dependent AQE of Bi₅O₇Br-NT. Reproduced from ref. 87 with permission from Wiley-VCH.

with a diameter of 20 nm ± 1.4 nm were uniformly distributed on the surface of TiO₂-OV nanosheets. The UV-vis spectrum of Au/TiO₂-OV showed absorption characteristics of Au loaded TiO₂ (Au/TiO₂) and TiO₂-OV, displaying a broad localized



surface plasmon resonance (LSPR) peak at 550 nm and an obvious absorption tail (Fig. 11b). Therefore, in the Au/TiO₂-OV hybrid, OVs on the surface of TiO₂ nanosheets chemisorbed and activated N₂ molecules, which were reduced to NH₃ by plasmon-induced hot electrons from Au nanoparticles in the presence of methanol as a hole scavenger (Fig. 11c). As a result, the optimal Au/TiO₂-OV exhibited an NH₃ photoproduction rate of 78.6 μmol g⁻¹ h⁻¹, which was 98 and 35 times higher than that of Au/TiO₂ and TiO₂-OV respectively (Fig. 11d). Recently, we reported a hybrid of 2D layered Ti₃C₂ and P25 TiO₂ (Ti₃C₂-P25), in which Ti₃C₂ served as the cocatalyst to facilitate charge separation due to its excellent electrical conductivity. The DFT calculation (Fig. 11e) revealed that the N₂ adsorption energies were ranked as Ti₃C₂ (2.731 eV) > VOs-TiO₂ (0.342 eV) > TiO₂

(0.170 eV), implying that Ti₃C₂ had a much stronger N₂ chemisorption effect compared with VOs-TiO₂ and TiO₂. Therefore the introduction of Ti₃C₂ on P25 could accelerate N₂ chemisorption, leading to five times higher photocatalytic yield of NH₃ than that on pure P25. More importantly, water acted as the electron donor in this study and stoichiometric O₂ production was also observed, which was one of a few examples that could observe the stoichiometric reduction and oxidation products.¹⁰³

In addition to TiO₂-based photocatalysts, Pt-loaded ZnO was reported as an efficient photocatalyst for N₂ photoreduction which exhibited an NH₃ yield rate of 860 mmol g⁻¹ h⁻¹.¹⁰⁴ More recently, O-doped 1T-MoS₂ nanosheets with a large amount of SVs (SV-1T-MoS₂) were used as cocatalysts over CdS nanorods for

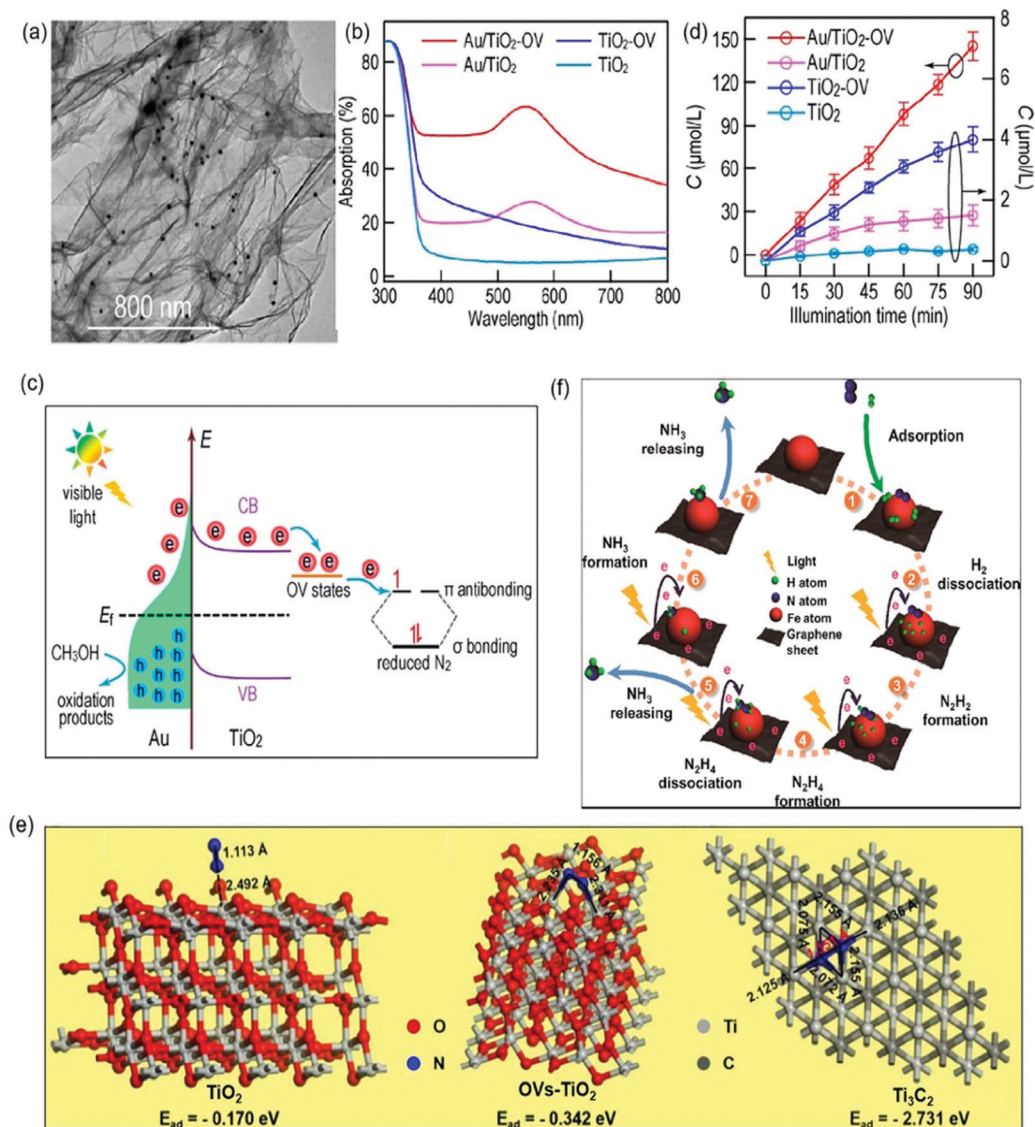


Fig. 11 (a) TEM image of Au/TiO₂-OV. (b) Adsorption spectra of different photocatalysts. (c) The proposed schematic for photocatalytic N₂ reduction over the Au/TiO₂-OV catalyst under visible-light irradiation. (d) Photocatalytic NH₃ production with different photocatalysts under visible light irradiation. Reproduced from ref. 96 with permission from ACS Publication. (e) Schematic of the adsorption structure on TiO₂ (A), OV-TiO₂ (B) and Ti₃C₂ MXenes (C). Reproduced from ref. 103 with permission from Elsevier Publication. (f) Schematic for photocatalytic N₂ reduction over Fe@graphene photocatalysts. Reproduced from ref. 108 with permission from ACS Publication.



photocatalytic N_2 reduction.¹⁰⁵ The DFT calculations revealed that the SVs and the metal properties of 1T-MoS₂ could effectively promote the separation of h^+/e^- and the presence of SV-1T-MoS₂ could also provide abundant active sites, resulting in enhanced N_2 adsorption and immobilization. Accordingly, a remarkable performance over the optimal SV-1T-MoS₂/CdS was realized with a superior NH_3 yield rate of $8220.83 \mu\text{mol L}^{-1} \text{h}^{-1} \text{g}^{-1}$ under simulated solar light irradiation. Besides, NiS was also studied as a cocatalyst over CdS nanorods for ammonia photosynthesis.¹⁰⁶ The DFT calculation indicated that the NiS cocatalyst could reduce the adsorption energy of N_2 from 1.16 eV on the CdS surface to -0.55 eV on the Ni-doped CdS surface. Further characterization revealed that charge separation could also be enhanced by loading NiS on the CdS surface. Because of these properties, a favorable NH_3 yield of 2.8 mg L^{-1} was reported within the first hour in pure water under full-spectrum irradiation. Recently, single-atom Pt modified triazine framework (CTF) nanosheets (Pt-SACs/CTF) were developed for photocatalytic NH_3 production.¹⁰⁷ It was found that the Pt-N₃ sites were formed in Pt-SACs/CTF, leading to more negative conduction band position with accelerated thermodynamics and fast interfacial charge migration ability, and therefore a high NH_3 production rate of $171.4 \mu\text{mol g}^{-1} \text{h}^{-1}$ under visible light irradiation. Most recently, we synthesized a ternary Ru/RuO₂/g-C₃N₄ system for N_2 photoreduction, which exhibited a photocatalytic NH_3 yield of $13.3 \mu\text{mol g}^{-1} \text{h}^{-1}$ whilst no NH_3 could be measured for pure g-C₃N₄ under the same condition. Such significant improvement in photocatalytic NH_3 synthesis for Ru/RuO₂/g-C₃N₄ was not only because of Ru and RuO₂ acting as cocatalysts to promote electron and hole transfer respectively, but also due to the advantage of Ru for N_2

chemisorption and activation.¹⁰⁸ Besides that, carbon-based materials especially graphene are also often used as cocatalysts to prevent the photoinduced electron-hole recombination and promote the electron transfer due to their prominent electrical conductivity.^{109,110} For instance, a Fe modified three-dimensional graphene (Fe@graphene) photocatalyst was reported for NH_3 production.¹¹⁰ In the Fe@graphene system (Fig. 11f), graphene generated hot electrons under light irradiation and transferred those electrons to Fe, while Fe acted as an electron sink and provided the catalytic active sites for the adsorption and activation of N_2 .

Junction structure. The combination of two semiconductors as a photocatalyst would overcome the shortages of each component, and construct the heterostructure or Z-scheme system with improved photoelectric properties for broad applications.^{111–115} In one work, Bi₂MoO₆ was implanted on g-C₃N₄ nanosheets to enhance the photocatalytic N_2 reduction.¹¹¹ In this case, the obtained g-C₃N₄/Bi₂MoO₆ heterojunction showed effective separation of charge carriers and enhanced photocatalytic performance for NH_3 production. In another report, MXene-derived TiO₂@C incorporated with g-C₃N₄ (TiO₂@C/g-C₃N₄, Fig. 12a) was prepared for efficient NH_3 photocatalysis.¹¹² Compared with the pristine g-C₃N₄, TiO₂@C/g-C₃N₄ showed a significantly decreased PL intensity (Fig. 12b), which indicated that the heterostructure promoted the charge transfer, which in turn inhibited the recombination of photoinduced electron-hole pairs. Furthermore, the transient photocurrent density of TiO₂@C/g-C₃N₄ was about 1.7 times higher than that of g-C₃N₄ (Fig. 12c), suggesting the improved separation efficiency of the charge carriers. The N_2 -TDP result revealed that the Ti³⁺ species on TiO₂ and carbonaceous components were responsible

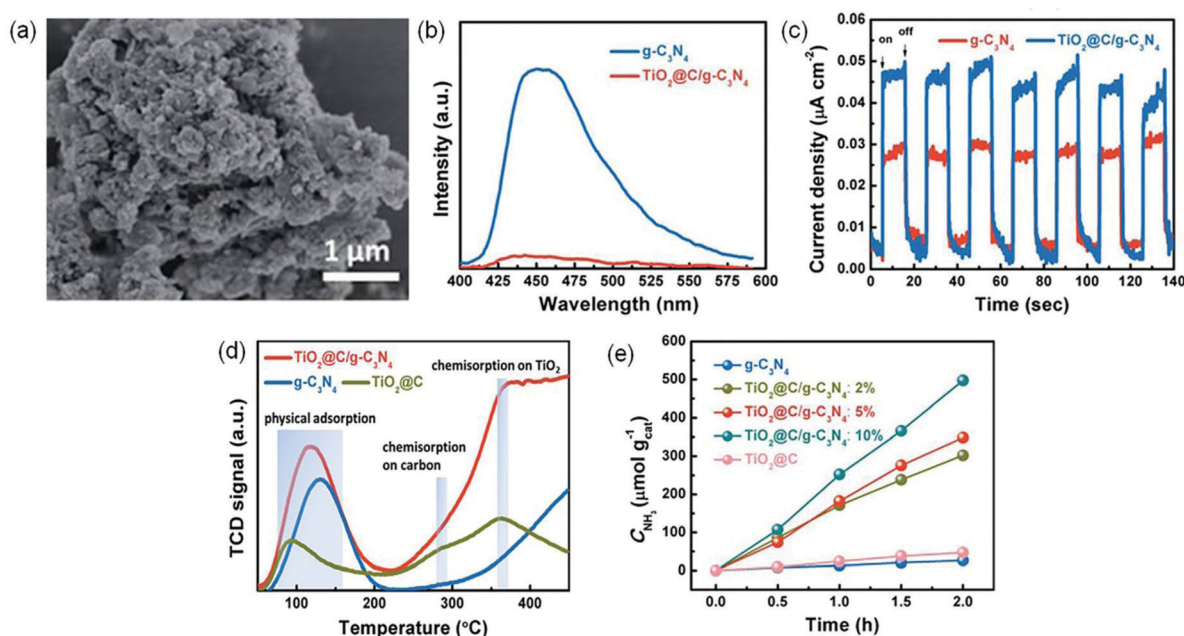


Fig. 12 (a) SEM image of TiO₂@C/g-C₃N₄. (b) PL spectra of TiO₂@C/g-C₃N₄ and g-C₃N₄. (c) Transient photocurrent responses of TiO₂@C/g-C₃N₄ and g-C₃N₄. (d) The N_2 -TDP profiles. (e) The yield of produced NH_3 under visible light irradiation ($\lambda > 420 \text{ nm}$). Reproduced from ref. 112 with permission from The Royal Society of Chemistry.



for the adsorption and activation of N_2 in $TiO_2@C/g-C_3N_4$ (Fig. 12d). As a result, a superior performance over the optimal $TiO_2@C/g-C_3N_4:10$ (10 represents the molar ratio of starting materials of melamine/ Ti_3C_2) was realized with an NH_3 production yield of $250.6 \mu\text{mol g}^{-1} \text{h}^{-1}$ using methanol as a scavenger under visible light irradiation (Fig. 12e). Recently, a p-n junction of $Bi_2MoO_6/OV-BiOBr$ combining n-type Bi_2MoO_6 nanorods with OV-rich p-type $BiOBr$ nanosheets displayed efficient charge separation and a broad range of light absorption.¹¹⁶ Moreover, OVs on the surface of $Bi_2MoO_6/OV-BiOBr$ were beneficial for the chemisorption and activation of N_2 , and the nanosheet morphology could provide a large specific surface area with more active sites for photoreduction of N_2 . Benefiting from these advantages, the optimal $Bi_2MoO_6/OV-BiOBr$ exhibited a photocatalytic yield of NH_3 of $81.0 \mu\text{mol g}^{-1} \text{h}^{-1}$ under visible light illumination in pure water, and no O_2 production was reported here.

Conclusion and perspectives

The photosynthesis of NH_3 is an energy-saving process and has high potential for substantial contribution to economic and social sustainability. Compared with the traditional Haber-Bosch process and the emerging electrochemical process, photocatalytic NH_3 synthesis exhibits very low efficiency, mostly at micromoles $g^{-1} \text{h}^{-1}$ if using water as the electron donor, and thus is far from practical application. However as photocatalysis can be driven by abundant solar energy, it is a carbon-free process. Besides, it requires a solar energy input of $208.3 \text{ MJ kg}^{-1}\text{-}NH_3$ compared with that of $339.1 \text{ MJ kg}^{-1}\text{-}NH_3$ in the electrochemical process.¹¹⁷ Thus, it has strong potential to achieve NH_3 synthesis with affordable costs in particular taking into account the sustainability and the trend of decarbonization in the industry. However there is a long way to go in order to realize this potential.

This review does not detail different photocatalysts as there are a few reviews on them, but analyses surface engineering and interface engineering in photocatalysts developed recently for the photocatalytic NH_3 synthesis, involving oxygen vacancies, nitrogen vacancies, carbon vacancies, sulfur vacancies, metal and non-metal doping, facet modulation, nanostructure engineering, and heterostructure construction such as loading with cocatalysts and modification with other semiconductors, together with the mechanism of N_2 reduction reaction and the reliable means for the detection of NH_3 .

Great progress has been made in the study of photoreduction of N_2 to NH_3 as briefed in Fig. 13, but there are still many problems, some of which are very critical. First, the majority of the studies so far reported use an efficient and expensive hole scavenger to get long lived photoelectrons for N_2 reduction. However, a practical pathway should use water as the only electron donor and to provide protons for the photosynthesis of NH_3 , as such NH_3 synthesis by photocatalysis is scientifically meaningful and economically sustainable. The slow progress linked to this profile is mainly due to the very limited understanding in this field. Second, there are few reports specifying

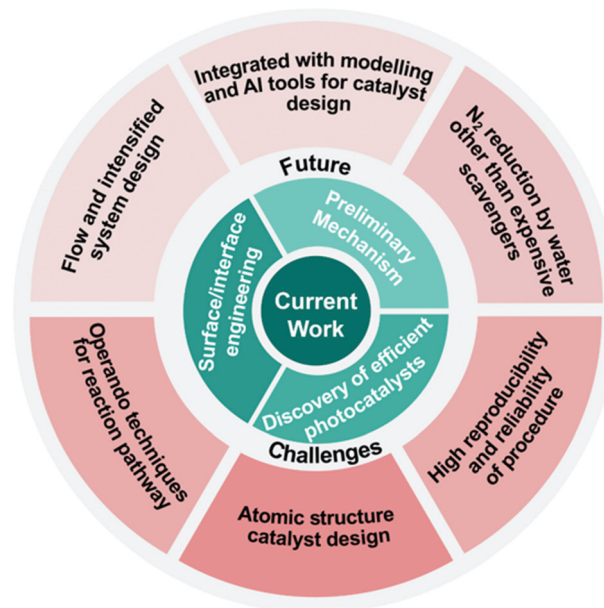


Fig. 13 Current work and future challenges highlighted in this review.

the stoichiometric chemistry in the photochemical process. Without the stoichiometric O_2 gas produced, such N_2 reduction to NH_3 is questionable. Furthermore, as O_2 could be leaked from air, the isotopic measurement of O_2 production is necessary. In addition, since H_2 production would compete with NH_3 synthesis, a comprehensive study should detail the selectivity between the two products and discuss the strategy to favor NH_3 synthesis. In addition an experimental error bar is very crucial for such NH_3 synthesis in order to improve the confidence as almost all studies were carried out in a batch reactor with quite a small amount of NH_3 produced, which was quite easy to be interfered by occasional factors.

Apart from the issues mentioned above, photocatalytic NH_3 synthesis also faces a few tough challenges. The majority of the studies show that the NH_4 amount produced is less than $10 \mu\text{mol h}^{-1}$ (instead of the unit of $\mu\text{mol g}^{-1} \text{h}^{-1}$) when using water as the electron donor and 10 times enhancement is achieved when using an organic hole scavenger. This amount is lower than the NH_3 amount in the river water which is about 0.01 mM .¹¹⁸ Besides, since the NH_3 yield is extremely easy to be interfered by environmental impurities and different detection agents, it is difficult to obtain reliable and reproducible results by just one analysis method.¹¹⁹ At least two methods have to be applied to prove the consistent results and to improve the reproducibility of the process.

The surface states of a catalyst can dominate its performance, such as surface defects. However, to quantify the correlation of the surface defects with the catalytic activity is very challenging at present. This should be undertaken in the future in order to distinctly guide catalyst surface state control. Furthermore the cocatalyst is another key factor to control both charge separation and catalytic performance. The majority of the cocatalysts reported are either large particles or nanosized particles. Single atomic catalysis is an emerging area, which



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