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## Electron transfer within $\beta$ -diketiminato nickel bromide and cobaltocene redox couples activating $\text{CO}_2$ †‡

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**Reduction of  $\beta$ -diketiminato nickel(II) complexes ( $\text{L}^{\text{tBu}}\text{Ni}^{\text{II}}$ ) to the corresponding nickel(I) compounds does not require alkali metal compounds but can also be performed with the milder cobaltocenes.  $\text{L}^{\text{tBu}}\text{NiBr}$  and  $\text{Cp}_2\text{Co}$  have rather similar redox potentials, so that the equilibrium with the corresponding electron transfer compound [ $\text{L}^{\text{tBu}}\text{Ni}^{\text{I}}\text{Br}$ ][ $\text{Cp}_2\text{Co}^{\text{III}}$ ] (ETC) clearly lies on the side of the starting materials. Still, the ETC portion can be used to activate  $\text{CO}_2$  yielding a mononuclear nickel(II) carbonate complex and ETC can be isolated almost quantitatively from the solutions through crystallisation. The more negative reduction potential of  $\text{Cp}^*\text{Co}$  shifts the equilibrium formed with  $\text{L}^{\text{tBu}}\text{NiBr}$  strongly towards the ETC and accordingly the reaction of such solutions with  $\text{CO}_2$  is much faster.**

Monovalent nickel has the reductive power to activate substrates by injecting an electron or by transferring electron density.<sup>1</sup> A couple of nickel enzymes are known or proposed to involve nickel(I) intermediates in developing functions like the conversion of carbon oxides.<sup>1</sup> Consistently, in the last decade the  $\beta$ -diketiminato nickel(I) moiety has demonstrated its potential to activate small molecules, such as  $\text{CO}_2$  and  $\text{CO}$ , but also  $\text{N}_2$ ,  $\text{H}_2$ ,  $\text{P}_4$ ,  $\text{SF}_6$ ,  $\text{O}_2$ , formate,  $\text{NO}$ , Azides and organonitro and -nitroso compounds.<sup>1–16</sup> Typically, it is generated by the reaction of a  $\beta$ -diketiminato nickel(II) bromido precursor with elemental sodium or potassium,  $\text{KC}_8$  or  $\text{Na/Hg}$ , and subsequently it has three possibilities to reach a stable coordinative saturation: (i) it binds a solvent donor molecule or a donor from the gas phase, like  $\text{N}_2$ ; (ii) the alkali metal halide generated concomitantly remains in the coordination sphere; (iii) the aryl ring of a second

molecule is coordinated so that a dimerization occurs.<sup>4,9,17</sup> A further option becomes conceivable, if reductants are used that produce a weakly coordinating cation, namely, the preservation of the  $\beta$ -diketiminato nickel bromide core. We were interested to examine the behaviour of such a system, especially in a redox regime that is not as harsh as under conditions where alkali metals are used as reductants, bearing in mind that nickel enzymes, which reductively activate  $\text{CO}$  or  $\text{CO}_2$ , also work at comparatively mild potentials.

For the selection of suitable reductants, the half-wave potentials had to be considered. In THF solution  $\text{L}^{\text{tBu}}\text{NiBr}$  ( $\text{L}^{\text{tBu}} = [\text{HC}(\text{C}(\text{CMe}_3)\text{NC}_6\text{H}_3(\text{tPr})_2)_2]^-$ ) exhibits a reversible redox event at  $-1.37$  V (vs.  $\text{Fc}/\text{Fc}^+$ ) that corresponds to the  $\text{Ni}^{\text{II}}/\text{Ni}^{\text{I}}$  redox couple (see Fig. S3, ESI†). Contemplating the chemical conversion of the nickel(II) complex to a reduced product, the Nernst equation for reversible electron-transfer processes requires the reductant to have a redox potential that is 0.118 V more negative to achieve a yield of 91%. The formal potentials of  $\text{Cp}_2\text{Co}$  and  $\text{Cp}^*\text{Co}$  in THF are  $-1.33^{18}$  V and  $-1.85^{19}$  V (vs.  $\text{Fc}/\text{Fc}^+$ ), respectively. Hence, in the case of  $\text{Cp}^*\text{Co}$  an almost quantitative electron transfer was to be expected. However,  $\text{Cp}_2\text{Co}$  was anticipated to represent an interesting case, too, as the formal potentials of both reactants are almost at the same level, so that the corresponding redox equilibrium should be subject to subtle influences.

Combining toluene solutions of  $\text{L}^{\text{tBu}}\text{NiBr}$  and  $\text{Cp}^*\text{Co}$  led to red solutions, from which upon standing for 1 h a brown crystalline solid precipitated. Single crystal X-ray analysis combined with spectroscopic data revealed that it corresponded to the expected product [ $\text{L}^{\text{tBu}}\text{Ni}^{\text{I}}\text{Br}$ ][ $\text{Cp}^*\text{Co}^{\text{III}}$ ], **1**. The determined structure (Fig. 1) features a  $\text{L}^{\text{tBu}}\text{NiBr}$  beside a  $\text{Cp}^*\text{Co}$  entity, and the relevant bond lengths and angles indicate that the electron has been transferred from  $\text{Co}$  to  $\text{Ni}$ , *i.e.* that **1** corresponds to an ionic compound: The bond distances around the nickel centre of the [ $\text{L}^{\text{tBu}}\text{NiBr}$ ]<sup>−</sup> anion in **1** ( $d(\text{Ni}-\text{N}) = 1.895(4)$ ,  $1.904(4)$  Å;  $d(\text{Ni}-\text{Br}) = 2.3447(8)$  Å) are significantly longer than those found in the neutral complex  $\text{L}^{\text{tBu}}\text{NiBr}$  (see Fig. S16 (ESI†),

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† Dedicated to Professor Dr Wolfgang Kaim on occasion of his 70th birthday.

‡ Electronic supplementary information (ESI) available: Experimental section, additional figures and X-Ray crystallographic data. CCDC 2023970, 2023971, 2024643. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d0cc06983c







Fig. 4 Offset cyclic voltammograms of  $L^{tBu}NiBr$  (green),  $[L^{tBu}NiBr][Cp^*_2Co]$  (red) and  $[L^{tBu}NiBr][Cp_2Co]$  (black) in THF/0.1 M  $[tBu_4N][PF_6]$ , scan rate  $100\text{ mV s}^{-1}$ . Dotted line denotes  $Ni^{II}/Ni^I$  couple.

side. In the solid state – removed from the equilibrium –  $[L^{tBu}Ni^I Br][Cp_2Co^{III}]$  is stable as an ion-pair.

Naturally, it was of interest how such a solution behaves electrochemically in comparison to solutions of  $[L^{tBu}NiBr]$  and **1** (Fig. 4). While  $L^{tBu}Ni^{II}Br$  shows the abovementioned single  $Ni^{II}/Ni^I$  redox event at  $-1.37\text{ V}$  (process I), CV scans of **1** as expected show two further reversible events at more negative potentials. Process II is assigned to the  $Co^{III}/Co^{II}$  couple, while process III at  $-3.16\text{ V}$  is typical for  $Co^{II}/Co^I$  of  $Cp^*_2Co$ . Consequently, three events should be expected for **2**, too, but the lower electron donating effect of Cp with respect to  $Cp^*$  for the cobaltocene means here the  $Ni^{II}/Ni^I$  and  $Co^{III}/Co^{II}$  events coincidentally overlap ( $E_{1/2}^{(I)} = -1.37\text{ V}$ ), and the  $Co^{III/I}$  event occurs at a more positive potential ( $E_{1/2}^{(II)} = -2.55\text{ V}$ ).

The increased current response of process I with respect to process II (average  $i_p^{(I)}/i_p^{(II)} = 1.2$ ) is consistent with this interpretation.

Finally, bearing in mind the background outlined in the introduction, the potential of **1** and **2** to activate  $CO_2$  was investigated. The nickel(i) complex  $[L^{tBu}Ni^I-N_2-Ni^I L^{tBu}]$  containing a labile dinitrogen ligand had been found to reductively couple  $CO_2$  to give oxalate.<sup>2</sup> After double reduction of  $[L^{tBu}Ni^I-N_2-Ni^I L^{tBu}]$  with two equivalents of potassium the resulting  $[L^{tBu}Ni^I-N_2-Ni^I L^{tBu}]K_2$  had proved capable of cleaving  $CO_2$  to give nickel-bound CO and a carbonate complex  $[L^{tBu}Ni^{II}-CO_3]K$ , which aggregates through interactions involving the potassium ions to yield a hexamer.<sup>2</sup> As in **1** and **2** the nickel complex anion contains only one reducing equivalent, a reaction analogous to  $[L^{tBu}Ni^I-N_2-Ni^I L^{tBu}]$  with formation of oxalate appeared intuitively more likely, but instead of the readily dissociating neutral  $N_2$  ligand more strongly bound anionic bromido ligands are found in **1** and **2**, so that the nickel centres are not similarly accessible. Indeed, this significantly altered the course of the reaction. Upon addition of  $CO_2$  to solutions of **1** in  $C_6D_6$  or toluene, within a minute the mixture became cloudy and a microcrystalline solid precipitated. This was analyzed by



Fig. 5 Ball and stick drawing of the solid state structure of the  $[L^{tBu}NiCO_3]^-$  anion in **3**.

means of ATR-IR spectroscopy and while no  $Ni-CO^3$  or oxalate<sup>2</sup> species were detected, bands between  $1550$  and  $1650\text{ cm}^{-1}$  were detected, indicating the formation of a carbonate product (further assignment is difficult due to band overlap). In contrast, reaction of **2** with  $CO_2$  took much longer and thus only after 30 minutes the solution started to become cloudy, which is reasonable, as **2** can only react out of an equilibrium, in which it exists as a minor component. However, the slow reaction rate led to the formation of a crystalline solid after storage for one week and a crystal structure analysis indeed revealed the formation of the mononuclear carbonate complex  $[L^{tBu}NiCO_3][Cp_2Co]$ , **3** (Fig. 5). Extensive efforts to grow high quality crystals remained without avail, so that the data quality does not permit a discussion of metric parameters, but the molecular structure determined and the constitution are without doubt. A carbonate anion is symmetrically coordinating the nickel centre in a square planar fashion, so that the structure of **3** is rather different as compared to the one of  $\{[L^{tBu}Ni^{II}-CO_3]K\}_6$ , where the  $K^+$  counterions actively participate in the structure construction and lead to aggregation to a hexanuclear complex.<sup>2,24,25</sup> A similar outcome as in case of the reaction between **1** and  $CO_2$  was also revealed for **2** by ATR-IR spectroscopy (bands between  $1550$  and  $1650\text{ cm}^{-1}$ ) and also in this case no  $Ni-CO^3$  or oxalate<sup>2</sup> complexes were detected.

Formation of **3** requires cooperation of two equivalents of **2**, which activate one  $CO_2$  molecule, likely yielding in a  $Ni^{II}-CO_2^{2-}$  intermediate, which can react with a second molecule of  $CO_2$  to yield CO and carbonate (Scheme 1).<sup>26</sup> Hence, concomitantly one equivalent of  $[L^{tBu}NiBr]$  should be produced as well as one equivalent of  $[Cp_2Co]Br$ , which, however, was not isolated. Similar cooperativity of two  $Ni^I$  centres in  $CO_2$  activation has been reported utilising different PNP-pincer ligands ( $L^{PNP}$ ). While in case of a rigid ligand the resulting  $L^{PNP}Ni^{II}-CO_2^{2-}-Ni^{II}L^{PNP}$  appeared stable in contact with excessive  $CO_2$ ,<sup>27</sup> utilising a more flexible ligand did also lead to concomitant reaction with an additional molecule of  $CO_2$ , forming a dinuclear  $L^{PNP}Ni^{II}-CO_3-Ni^{II}L^{PNP}$  unit and CO.<sup>28</sup> However, no examples have been reported so far where a cooperative activation of  $CO_2$  leads to a mononuclear nickel complex, and thus our system is closing this gap. We hypothesise that the presence of the bromido ligand, which remained in the coordination sphere of the reduced nickel centre due to the employment of cobaltocenes as reducing agents (converting into non-coordinating



