


 Cite this: *Chem. Commun.*, 2018, 54, 11284

 Received 11th July 2018,
 Accepted 18th September 2018

DOI: 10.1039/c8cc05591b

rsc.li/chemcomm

The semiquinone radical anion of 1,10-phenanthroline-5,6-dione: synthesis and rare earth coordination chemistry†

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Reduction of 1,10-phenanthroline-5,6-dione (pd) with CoCp^R₂ resulted in the first molecular compounds of the pd^{•−} semi-quinone radical anion, [CoCp^R₂]⁺[pd]^{•−} (R = H, (1); R = Me₄, (2)). Furthermore compounds 1 and 2 were reacted with [Y(hfac)₃(thf)₂] (hfac = 1,1,1-5,5,5-hexafluoroacetylacetonate) to synthesise the rare earth-transition metal heterometallic compounds, [CoCp^R₂]⁺[Y(hfac)₃-(N,N'-pd)]^{•−} (R = H, (3); R = Me₄, (4)).

Electron transfer reactions are fundamental to chemical processes. Stable organic and organometallic radicals are widely utilised, particularly in small molecule activation and catalysis.¹ Redox-active ligands can act as electron reservoirs to achieve two electron processes, such as oxidative addition or reductive elimination, that would otherwise be inaccessible. This is a successful strategy in transition metal chemistry for the replacement of precious metals with earth abundant metals² and has also been demonstrated for the f-elements.^{3,4}

The most common use of a f-element radical is the use of samarium diiodide in organic chemistry. However, organometallic complexes of the f-elements have proved particularly suited for the reductive activation and stabilisation in well-defined molecular complexes, of some of the most unusual and biologically relevant radical species *e.g.* CO^{•−}, CO₂^{•−} and NO^{•−}.^{5–7} The number of examples of f-elements complexed to redox-active ligands or radical anions is still modest but this is a growing area of research activity; for magnetism, multi-electron transfer chemistry and functional materials.^{4,8} Heterobimetallic complexes remain a particular synthetic challenge. In order to realise mixed-metal complexes, a bridging ligand with different coordination sites is needed so that metal ions can be selected and installed step-wise to acquire the desired properties. However, the

bridging ligands used in f-element chemistry have identical binding sites that lead near exclusively to homobimetallic complexes.

Phenanthroline dione (pd) is a commercially available yet highly underutilised redox-active ligand. It constitutes a three-membered electron transfer series, where the neutral form can be sequentially, one-electron reduced to give the monoanionic and dianionic forms (Scheme 1). The terminal members are diamagnetic, whereas the monoanionic form is a radical with an *S* = 1/2 ground state. The most prized features of this ligand are its two binding sites. In the dianionic form of the ligand (right, Scheme 1), the *N,N'*-pocket is a π -accepting diimine unit analogous to 2,2'-bipyridine, and distinct from the *O,O'*-pocket which is a π -donating diolate, akin to a catecholate.

There is literature precedent for the formation of multi-metallic complexes using the dianionic form of the ligand.^{9–12} We also recently reported heterometallic rare earth-transition metal complexes utilising pd as a template.¹³ Our two-step route was necessary because although pd is reported to react with a number of reducing agents, the products are insoluble in common solvents, precluding collection of spectroscopic data and definitive assignment of redox state, but more importantly further reactivity.^{11,14,15}

In contrast to the well-developed chemistry of transition metal semiquinones,¹⁶ there are only three reports of the *in situ* characterisation of pd^{•−}.^{17–19} These include two spectro-electrochemical studies in combination with EPR, and an EPR study.

Here we report the first organometallic synthons of the radical anion of pd, [CoCp^R₂]⁺[pd]^{•−} (R = H, (1); R = Me₄, (2)).



Scheme 1 Three-membered electron transfer series for pd showing calculated bond distances.

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† Electronic supplementary information (ESI) available. CCDC 1854232 and 1854233. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c8cc05591b





Fig. 2 Comparison of the X-band EPR spectrum of **1** (experimental conditions: frequency, 9.8461 GHz; power, 0.63 mW; modulation, 0.007 mT), and **3** at 293 K (experimental conditions: frequency, 9.8653 GHz; power, 0.63 mW; modulation, 0.02 mT) recorded in MeCN solution at 293 K. Experimental data are represented by the black lines and simulation depicted by the red traces.

temperature (Fig. 2). Spectral simulation of **1** was achieved with $g_{\text{iso}} = 2.0053$ consistent with a heterocyclic organic radical. The spectrum has rich hyperfine structure that was reproduced by including coupling to three pairs of protons in the *ortho*, *meta* and *para* positions of the aromatic ring (attached to C1–C3 in Fig. 1). These yielded couplings of 1.40, 1.08, and $1.56 \times 10^{-4} \text{ cm}^{-1}$, respectively. In addition, coupling to the ^{14}N ($I = 1$) nuclei was estimated at $0.51 \times 10^{-4} \text{ cm}^{-1}$. These spin-Hamiltonian parameters are very similar to previously reported spectra of $\text{pd}^{\bullet-}$.^{17–19} Coordination of the $\text{pd}^{\bullet-}$ to the Y^{III} d^0 ion results in a slightly wider spectrum with broader lines. The increase in line width is a consequence of slightly perturbed molecular tumbling commensurate with the inclusion of the bulky $[\text{Y}(\text{hfac})_3]$ moiety. The centre of the spectra shifts to $g_{\text{iso}} = 2.0041$ with a minor change in the magnitude of the ^1H hyperfine interaction to 1.41, 0.52, $1.56 \times 10^{-4} \text{ cm}^{-1}$ (*cf.* **1** above), and an invariant ^{14}N hyperfine splitting at $0.45 \times 10^{-4} \text{ cm}^{-1}$. Thus, the most profound change is the halving of the coupling to the proton in the *meta* position. The absence of ^{89}Y ($I = 1/2$) coupling is likely a result of the polarised bonding with this Lewis acidic metal as well as its small nuclear moment ($g_{\text{N}} = -0.275$) such that it is less than the experimental linewidth. Moreover, both spectra are consistent with a $\text{pd}^{\bullet-}$ radical where the spin is nearly exclusively localised to the *O,O'* end of the molecule as supported by computational data.

The structure of the anion in **1** was geometry optimised using the BP86 functional and is consistent with the radical

$\text{pd}^{\bullet-}$ ($S = 1/2$) form of the heterocycle (Scheme 1). The calculated C–O and C–C bond distances of the *O,O'*-pocket of 1.266 and 1.478 Å nicely match to the solid state structure (Fig. 1). In addition the C–C and C–N bond lengths exhibit a noticeable long-short-long pattern symptomatic of a quinoidal distortion commensurate with the addition of an unpaired electron to the aromatic system (Scheme 1).^{16,20} This effect diminishes the further the bond is from the *O,O'*-pocket. The Mulliken spin population analysis reveals 74% on the spin localised to this C_2O_2 unit, with a modest amount scattered over the phenanthroline part (Fig. S20, ESI†). The optimised structure of the complex ion in **3** depicts a square antiprism geometry adopted by the one $\text{pd}^{\bullet-}$ and three hfac^- ligands coordinated to the central Y^{III} ion (Fig. 3). The intraligand bond distances with the $\text{pd}^{\bullet-}$ are nearly identical to those observed for the free radical in **1**. Its attachment to the Y^{III} ion results in very little change in the structural change to the redox-active ligand. The Mulliken spin population shows an identical spin distribution for the coordinated $\text{pd}^{\bullet-}$ with the lion's share of the spin residing around the *O,O'*-pocket remote from the Y^{III} centre (Fig. 3). No spin density is found on the Y^{III} metal ion consistent with the weakly covalent bond diagnosed by EPR. These data are also consistent with *N,N'*-binding (*vide supra*) and in contrast to the perturbation to spin density and ^{89}Y coupling, expected to result from *O,O'*-binding of the Y^{III} centre.

Intensely coloured solutions of **1** and **3** are the result of a dominant absorption band at 558 and 639 nm, respectively. To gain insight into the nature of this transition, we have calculated the electronic spectra based on the electronic structure for the anion in **1** and **3**. Time-dependent (TD) DFT calculations were performed on the geometry optimised structures using MeCN as solvent in order to assign the electronic spectra. The vis-NIR spectra of both contain two transitions: 401 and 558 nm for **1**, and 364 and 639 nm for **3** (Fig. 4). These bands are well reproduced, in particular the intensity, using the TD-DFT protocol though are blue-shifted *ca.* 1200 cm^{-1} compared to the experiment. For **1**, the bands which are $\pi \rightarrow \pi^*$ transitions appear at 411 and 522 nm, result from HOMO–2 \rightarrow SOMO and

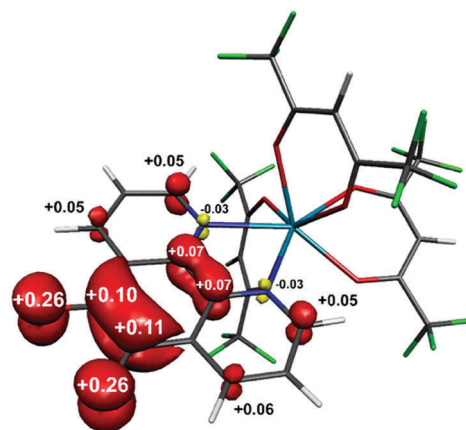


Fig. 3 Mulliken spin density distribution in $[\text{Y}(\text{hfac})_3(\text{pd})]^{\bullet-}$ (red: α -spin; yellow: β -spin).





Fig. 4 Overlay of the experimental (solid line) and calculated (dashed line) electronic absorption spectra for **1** (top) and **3** (bottom) recorded at ambient temperature in MeCN. Vertical bars represent individual calculated transitions. Inset orbitals depict the donor/acceptor molecular orbitals (MO) that constitute the band at 558 nm in **1** and 639 nm in **3**.

SOMO \rightarrow LUMO+1 excitations, respectively (Fig. S22, ESI[†]); the latter MOs are shown in the inset in Fig. 4. The analogous bands in **3** are calculated at 405 and 567 nm, respectively (Fig. 4). These are also assigned as $\pi \rightarrow \pi^*$ transitions within the $\text{pd}^{\bullet-}$, although in this instance as the organic radical is coordinated to a metal ion, they are specified as intraligand charge transfer (ILCT) transitions. Analysis of the donor and acceptor MOs shows the band at 567 is the SOMO \rightarrow LUMO+4 excitation, shown in the inset of Fig. 4. The higher energy band at 405 nm is predominantly a HOMO-2 \rightarrow LUMO transition (Fig. S23, ESI[†]); the order is different to **1** as it includes contributions from the $[\text{Y}(\text{hfac})_3]$ moiety. Again, the intensity of these two transitions is nicely reproduced by the TD-DFT calculations. These transitions are diagnostic of the singly-reduced radical $\text{pd}^{\bullet-}$,¹⁷⁻¹⁹ as also observed for related 9,10-phenanthrenesemiquinone ($\text{pq}^{\bullet-}$) complexes $[\text{MX}(\text{pq}^{\bullet-})(\text{PPh}_3)_2(\text{CO})]$ ($\text{M} = \text{Ru}$, $\text{X} = \text{Cl}$; $\text{M} = \text{Os}$, $\text{X} = \text{Br}$) where the dominant feature is also assigned as a $\pi \rightarrow \pi^*$ ILCT.^{21,22}

In conclusion, this work represents the discovery and success of a novel approach to the synthesis of rare earth-transition

metal complexes. The stability of the bridging $\text{pd}^{\bullet-}$ ligand, and the selectivity of N,N' -binding for $\text{Y}(\text{III})$ have been demonstrated. It is anticipated that compounds **1-4** will be useful synthons in pursuit of mixed-metal radical-bridged f-element complexes in the future.

The authors acknowledge the EPSRC (EP/L504786/1 and EP/M508056/1) and the University of Glasgow for funding. We also acknowledge the EPSRC UK National Crystallography Service at the University of Southampton.

Conflicts of interest

There are no conflicts to declare.

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