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Coumarin based Gold(I)-Alkynyl Complex: A new class of Supramolecular Hydrogelator

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Abstract

A phosphine-gold(I)-alkynyl-coumarin complex, [Au{7-(prop-2-ine-1-yloxy)-1-benzopyran-2-one}(DAPTA)] (**1**), was synthesized and the formation of long luminescent fibers in solution was characterized via Fluorescence Microscopy and Dynamic Light Scattering. The fibers presented strong blue and green luminescence, suggesting that the gold(I) in the complex increased intersystem crossing due to heavy atom effect, resulting in a significant increase in triplet emission. X-ray structure of the fibers indicates that both aurophilic, $\pi - \pi$ interactions and hydrogen bonding contribute to their formation in aqueous solvents.

Keywords: Gold(I)-Alkynyl Complexes; coumarin; luminescent gels.

Introduction

Supramolecular gels[1] have become a topic of increasing interest in recent years. Compared to traditional polymers, in which the individual monomers are linked irreversibly by covalent bonds, supramolecular linkages are reversible, which can result in new dynamic structures with interesting and new properties completely different from those observed on the individual molecules.[2]

In particular, self-assembled supramolecular hydrogels have attracted much attention during the past few decades because of promising high-technology applications in very diverse fields such as biomedicine, nanoelectronics and catalysis.[3-9] These colloidal materials can be formed either by low- or high-molecular-weight hydrogelators by means of non-covalent interactions (e.g., π–π stacking, van der Waals forces, hydrogen bonds, hydrophobic interactions).[10,11] Among these materials, metallogels[12-14] represent an important subclass where at least one metallic element has been incorporated into the gel network but they have been only a subject of study in the last few years. It should be highlighted the presence of metal ions since they may also modify the physical properties of the gels.[15]

In contrast to the development of coordination polymeric gelators, the use of discrete small-molecule metal complexes such as metallogelators has been relatively scarce and it was only recently that a number of reports on the gelation properties of metal complexes, such as Au(I), Au(III) and Pt(II), have appeared.[12,15-18] Interestingly, this process becomes unpredictable in some cases.

The formation of the supramolecular gels and metallogels depends on the concentration of the monomer and on the strength of the non-covalent interactions responsible to hold the monomers together.[2,19] In general, hydrogen bonds are among the most often used non-covalent interactions in the formation of supramolecular gels and polymers. Nevertheless, in the particular case of organometallic complexes like gold(I) derivatives, the possible establishment of Au(I)···Au(I) (aurophilic) bonds is also observed to be an important driving force responsible for the aggregation.

In fact, the formation of aurophilic interactions in the crystal structures of gold compounds, as well as in solution, has fascinated chemists for decades. The aggregation of gold compounds in this fashion arises from relativistic effects that are at a maximum

for gold compared to other heavy elements.[20-22] One reason for this interest is that gold compounds exhibit interesting luminescence properties, in both the solid-state and in solution that could be modulated by the presence of these Au···Au interactions.[22-27]

Aurophilic interactions can also lead to fascinating supramolecular architectures[21,22,28-32] and are known to present stabilization energies to their crystal structures in the range of 7-12 kcal/mol.[33,34] This energy range value should be stressed since it is comparable to that normally associated with conventional hydrogen bonding.

Among gold(I) supramolecular structures, gold(I) alkynyl derivatives are specially denoted due to the preference of gold(I) for a linear coordination geometry, together with the linearity of an acetylide unit and its π-unsaturated nature. These structural characteristics have made the alkynylgold(I) complexes attractive building blocks for organometallic oligomeric and polymeric materials which may possess unique properties in different fields.[22] Related to all of this, we have recently described the formation of luminescent hydrogelators with two gold(I) alkynyl derivatives with very simple chemical structure ([Au(4-pyridylethynyl)(phosph)] {phoshp = PTA (1,3,5-triaza-7-phosphaadamantane);[17] DAPTA = 3,7-diacetyl-1,3,7-triaza-5-phosphabicyclo[3.3.1]nonane)}.[18] In this work, the synthesis of a new hydrogelator that contains a propargyloxicoumarin group instead of pyridylethynyl is described (Chart 1). In order to try to understand better the gelation process with these discrete molecules, we report on the analysis of the aggregation process by different spectroscopical and microscopic techniques and the results compared with the X-ray crystal diffraction data retrieved for this complex.

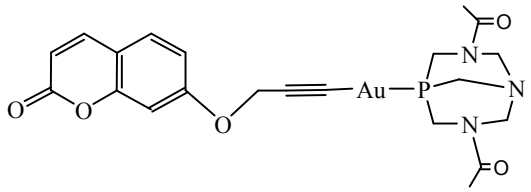

Chart 1

Materials and methods

Reagents

The synthesis of [Au{7-(prop-2-ine-1-yloxy)-1-benzopyran-2-one}(DAPTA)] (**1**) is described in previous reports.[35]

Physical Measurements.

Absorption spectra were recorded in a Varian Cary 5000 UV-Vis Spectrophotometer in a 3mL (10x10 mm wide) Quartz cell. Fluorescence measurements were performed in a Horiba-Jobin-Yvon SPEX Fluorolog 3.22, in a 3mL (10x10 mm wide) Quartz cell. Emission quantum yields were measured by employing 7-methoxy-4-methylcoumarin as a reference (ϕ_F = 0.124, 25 °C, methanol).

Dynamic Light Scattering experiments were performed in a Horiba Scientific Nanoparticle Analyzer SZ-100, in 3mL (10x10 mm wide) quartz cells. The samples were filtered before analysis using a 0.46 μm polystyrene membrane disc filter.

Fluorescence microscopy was recorded on an Axioplan 2ie Zeiss imaging microscope equipped with a NikonDXM1200F digital camera. Excitation light for fluorescence imaging was selected using filters in the range 300-400 and 450-490 nm respectively.

Preparation of the gel

Solid **1** was dissolved in water in different concentrations range (0.01 - 0.02% weight). The sample was maintained under sonication for *ca*. 20 minutes and left at room temperature at least for 4 hours.

X- ray crystal structure determination.

Data for **1** was collected at 123 K on an Agilent SuperNova diffractometer with Atlas detector using mirror-monochromated Cu-Kα radiation (λ = 1.54184 Å). The *CRYSALISPRO*[36] program was used for the data collection and processing of both crystals. The intensities were corrected for absorption using the multi-scan absorption correction method.[36] The structure of **1** was solved using direct methods with SHELXS solution program.[37] The structure was refined by full-matrix least-squares calculations based on F^2 using *SHELXL-2013*[38] integrated in the *WINGX*[39] program package. Hydrogen atoms were included in calculated positions as riding atoms, with *SHELXL-*

2013[38] defaults. X-ray crystallographic data for **1** is given in Table S1 (see Electronic supplementary information).

CCDC-1026387 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

Results and discussion

The synthesis and characterization of **1** was recently reported by us and its photophysical properties were reported in methanol.[35] Herein we describe the observed formation of hydrogels (see Figure S1) and compared with the previous data obtained with pyridylethynyl derivatives.[17,18] In the case of **1**, apart from the Au(I) atoms, the structure of the complex possesses a phosphine and a coumarin unit at both ends of the molecule. While the former moiety provides an increase in hydrophilicity, the latter moiety presents strong absorption and emission in the UV region of the spectrum. Thus, we measured the optical properties of the coumarin as a means to study the formation of aggregates which ultimately lead to new luminescent soft materials such as hydrogels.

Absorption/Fluorescence spectroscopy

The effect of concentration on the absorption and emission of complex **1** was primarily studied in methanol (Fig. 1) where aggregates are not expected to be formed.

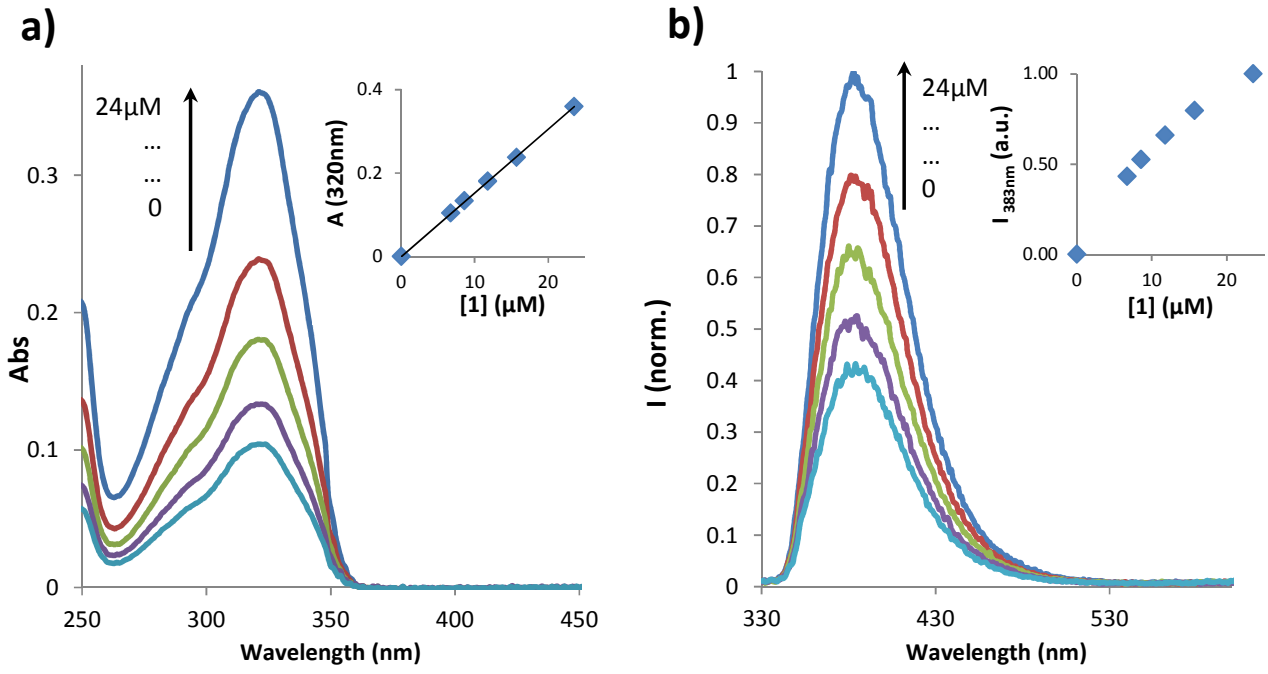

Figure 1. (a) Absorption spectra at different concentrations of **1** in methanol. Inset: Absorbance at 320 nm vs. concentration of **1**; (b) Fluorescence spectra at different concentrations of **1** in methanol. Inset: Intensity (I) at 383 nm versus concentration of **1**. λ_{exc} = 320 nm.

As can be seen in Fig. 1a and its inset, no changes were observed both in the shape as well as in the values for absorbance maximum of the spectra up to 24 μM, with a calculated molar absorptivitiy (ε) of 15259 M^{-1} cm^{-1} at 320 nm. Although the spectra show that linearity for the fluorescence/concentration ratio is only achieved up to 12 μM (the slight curvature in emission at 383 nm arises from inner filter and reabsorption effects), the calculated quantum yields remain constant throughout the entire studied concentration range (see Table 1).

Table 1. Fluorescence quantum yields of **1** in methanol at different concentrations.

[**1**] (μM)	ϕ_F
24	0.053
16	0.057
12	0.052
9	0.057
7	0.054

In water, complex **1** did not dissolve completely at concentrations above 100 μM. In fact, absorption spectra of **1** for concentrations ranging from 5 to 100 μM revealed an increase in the baseline which is a clear indication of precipitation/aggregation phenomena. The analysis of the absorption changes against concentration led us to calculate the critical gelation concentration as 28.5 μM (Figure S2).

The influence of temperature in the aggregation of **1** was also studied by recording the spectral changes in absorption between 25 and 80 ºC (Fig. 2).

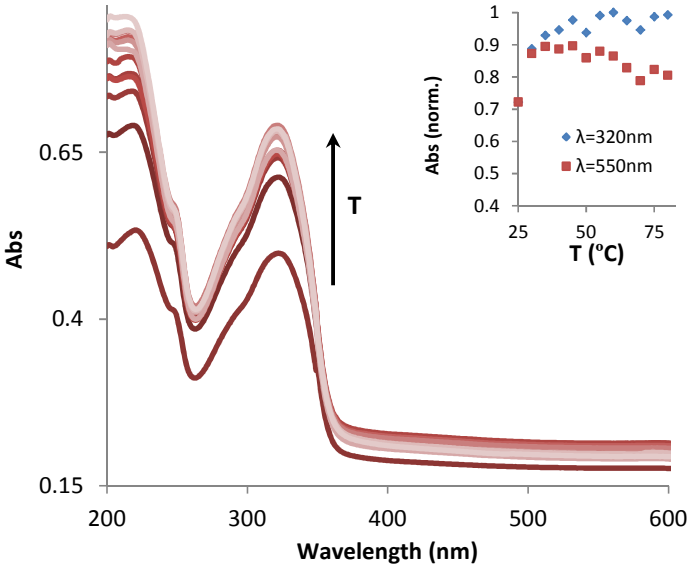

Figure 2. Absorption spectra of **1** (38 μM in water) at different temperatures.

Although the shape and the maximum absorption wavelength (320 nm) for the coumarin unit were constant, the value of absorbance at the maximum increased with temperature until reaching a plateau, as depicted in the inset of Figure 2. However, the values of absorbance at 550 nm, *i.e*. the tail of the spectrum corresponding essentially to the baseline, reveal the same trend. This rise of absorbance in baseline is typical for the increased opacity/optical density of the solution corresponding to precipitation of particles. Therefore, we assume that increased temperature induces aggregation of **1**. This behavior occurs mainly in the temperature interval 25-40 ºC, similar to NIPAM type polymers which collapse with temperature due to breaking of hydrogen bonds with water (solvent) molecules that simultaneously favors hydrophobic interactions.[40]

An hypothesis to explain these results can be found in the thermodynamics of the system. In most cases, the thermodynamics of aggregation can be described as:

$$\Delta G_{\text{aggregation}} = \Delta H - T\,\Delta S$$

where both ΔH and ΔS are negative, i.e., increasing temperature inhibits aggregation.

In this particular case, however, the molecule comprises a hydrophobic part (coumarin) and a hydrophilic moiety (phosphine), the initial formation of micellar-type structures is a possible aggregation mechanism (further evidence was obtained in the following sections). The hydrophobic parts are then surrounded by a shell of water molecules organized in a way to shade it from the solvent. For aggregation to happen, this shell must be displaced momentarily, which creates a rise in entropy and thus changes the initial formula towards this modified thermodynamic equilibrium:

$$\Delta G_{\text{aggregation}} = \Delta H - T\,(\Delta S_{aggregation} + \Delta S_{solvent})$$

If the solvent contribution (positive) is bigger than the aggregation contribution (negative) to entropy, then increasing the temperature will lead to a faster aggregation by increasing the driving force.

Dynamic Light Scattering

In order to obtain more insight on the size of the aggregates, screening tests at different concentrations were performed using Dynamic Light Scattering. The overall results show that aggregates can be found down to the micromolar concentration range in water (Figure S3). The size distribution of the aggregates allows the identification of a single population with a well-defined mean diameter (ranging from 70-140 nm, depending on the concentration). However, in all cases, a high polydispersity index is recorded, which is a strong indication for the presence of larger aggregates. Nevertheless, at the lowest concentrations (50µM and lower) and after filtration with a 0.46µm filter, we can obtain good mono-exponential fittings of the auto-correlation curves, with low polydispersity index (PDI) values, suggesting the presence of spherical nanostructures.

The influence of temperature on the formation/destruction of aggregates was studied by performing temperature ramps at 38 µM in water, as a means to complement the results of absorption experiments (Fig. 3).

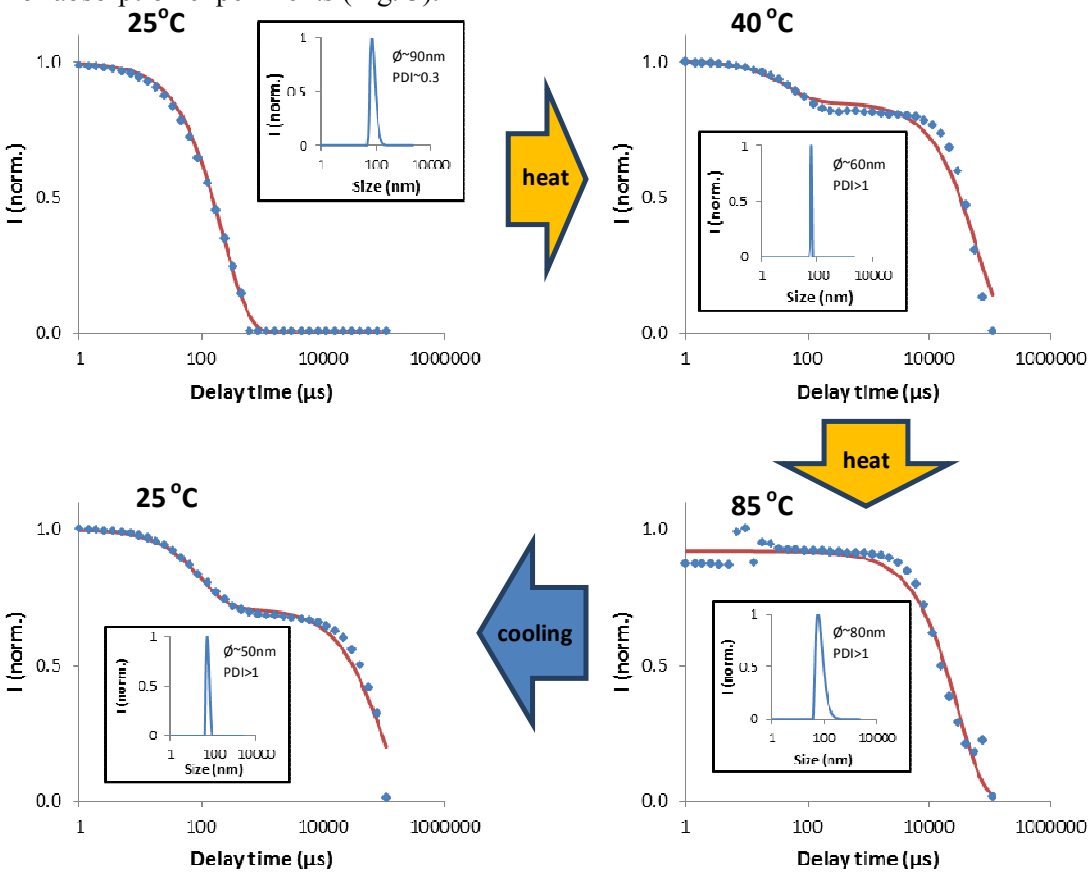

Figure 3. Evolution of DLS scans upon increasing the temperature from 25 ºC to 85 ºC and cooling back to 25 ºC.

As one can observe, a mono-exponential fitting of the DLS correlation function is achieved, with a monodisperse population of aggregates of about 90 nm, typical for micellar-like aggregates that could be created from hydrophobic interactions, as mentioned in the previous section. Raising the temperature to about 55 ºC leads to a bi-exponential function due to the appearance of a long decay signal, a clear indication for the formation of bigger structures.

This fact can be explained based on the molecular structure of complex **1**, if we consider that linear Au(I)-alkynyl complexes are known to form intermolecular aurophilic bonds and produce gels.[17,18] If this is the case, the formation of linear aggregates/fibers may explain the high PDI values, since the DLS models for determination of hydrodynamic assume a spherical shape for the analyzed samples.

Upon further increasing the temperature, we can no longer observe the initial population of smaller aggregates (nm range), although one cannot exclude their presence since the light scattering phenomenon is intrinsically linked to the size in a non-linear way, *i.e.,* bigger particles/structures scatter more light than smaller ones. Finally, by decreasing the temperature, the initial signal for small particles is partially recovered. These results are consistent with the aforementioned hypothesis for temperature-favoured aggregation.

Fluorescence microscopy

Images of the aggregates at different concentrations of **1** were obtained upon drying the solutions and analyzing the samples under the microscope. Figure 4 shows bright field images of dried solutions of **1**.

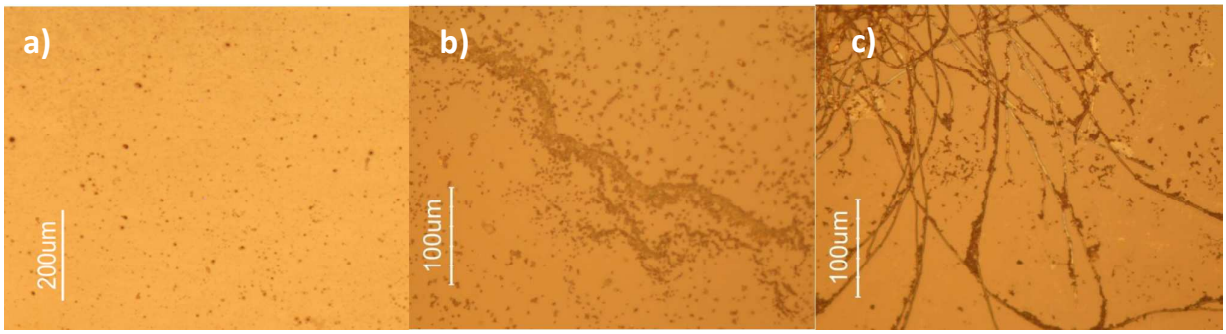

Figure 4. Bright field microscopic images of dried solutions of **1** at (a) 3.8 µM, (b) 38 µM and (c) 1000 µM.

Detailed observation of pictures from Fig.4 reveals the evolution of aggregation upon increasing concentrations of **1**. Indeed, at low concentration, only small spherical aggregates are observed which is consistent with the results from DLS, even though the size scales are quite different, since the solutions for DLS where previously filtered with a 460 nm filter and particles smaller than 460 nm are not seen under the microscope. A ten-fold more concentrated sample of **1** (image B) already reveals the presence of agglomerates of particles in a linear fashion, while at 1mM a highly reticulated network of fibers can be clearly identified in the picture which accounts for the observed high polydispersity of the samples in DLS measurements.

Pictures were also taken under UV irradiation using two different cut-off filters. Detailed analysis of these pictures shows that, while green emission is present throughout the whole reticulated network of fibers (Fig. 5b), blue emission can only be detected at the extremities (Fig. 5a).

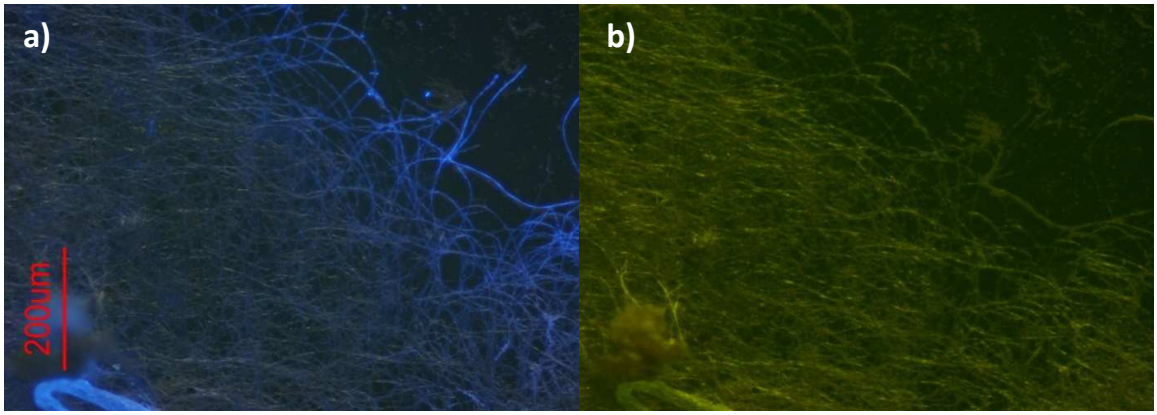

Figure 5. Fluorescence microscopy images of dried solutions of **1** at 1mM using a cut-off filter of (a) 300-400nm and (b) 450-490nm.

A possible explanation for this behavior is the heavy-atom effect induced by the gold atoms, *i.e.*, in the bulk of the aggregate the fibers are highly crosslinked, which causes the gold atoms and coumarins to be closer to each other than in the case of isolated linear fibers that are linked by the aurophilic interactions.

Indeed, since it has been shown that this kind of compound shows a linear architecture L1-Au-L2 and that it aggregates by Au···Au interactions,[22] the coumarin and gold atoms can never be closer than the intramolecular distance, which is in the case of **1** consists of four chemical bonds due to the methoxy-alkynyl spacer. However, once this "dense" aggregation phase is reached, gold and coumarin units can be, in average, at a shorter distance than the intramolecular distance, thus allowing for a more efficient intersystem crossing to take place and resulting in coumarin phosphorescence emission. Previous data obtained with ([Au(4-pyridylethynyl)(phosph)] {phoshp = PTA (1,3,5-triaza-7-phosphaadamantane)[17] and DAPTA = 3,7-diacetyl-1,3,7-triaza-5-phosphabicyclo[3.3.1]nonane)}[18] complexes have been a good reference for the analysis of these gels. It must be highlighted that the comparison with previous complexes led us to suggest that the presence of DAPTA phosphine significantly increases the entanglement of the aurophilic gels.

X-Ray crystallography

As mentioned above in the text, formation of particles can occur if we consider the amphiphilic character of **1** that yields micelle-type structure. The growth of fibers at higher concentrations may have resulted from the formation of weak aurophilic interactions in complex **1** between linearly coordinated Au(I) molecules.

Although, the formation of large aggregates with our gold(I) complexes has been investigated by different spectroscopic and microscopic techniques, X-ray crystallography is one basic powerful tool to visualize and analyze the structural features of metal complexes for the presence of intra- and intermolecular interactions. However, highlighting the demanding work with gels, we successfully isolated the single crystals of complex **1** suitable for X-ray analysis by diffusing hexane into a dichloromethane solution of **1**.

Selected bond distances and angles are summarized in Table 2.

Table 2. Selected bond lengths (Å) and angles (º) for **1**.

Distance	(Å)	Angle	(º)
Au1-Au2	3.1274(4)	Au1-Au2-Au3	147.611(12)
Au2-Au3	3.1316(4)	C22-Au1-P2	176.9(2)
Au1-Au3	6.0106(5)	C1-Au2-P1	178.2(2)
Au1-P2	2.2782(18)	C43-Au3-P3	177.9(2)
Au2-P1	2.2706(18)	C1-C2-C3	175.3(13)
Au3-P3	2.2647(19)	C22-C23-C24	178.4(10)
Au1-C22	2.043(8)	C43-C44-C45	175.9(9)
Au2-C1	2.013(7)	P1-Au2-Au1	90.52(5)
Au3-C43	2.013(6)	P1-Au2-Au3	91.44(5)
C1-C2	1.185(12)	P3-Au3-Au2	90.50(5)
C2-C3	1.450(13)	P2 Au1 Au2	92.40(5)
C10-C11	1.463(15)		

C31-C32	1.438(15)
C52-C53	1.418(15)

Crystal structure of **1** confirms the coordination of alkyne and phosphine ligands to gold, as shown in Figure 6.The asymmetric unit solved in the triclinic space group *P*-1, contains three independent alkyne and phosphine ligands coordinated to gold(I) displaying typical linear geometry with C−Au−P angles ranging from 176.9(2)° to 178.2(2)°. These values are similar to those of recently reported gold alkynyl derivative containing PTA phosphine ligands.[41]

As seen in Table 1, shortened C-C single bond distances are found next to carbonyl oxygen for the organic ligand[35], due to the conjugation of π electrons of the aromatic rings and with those of the carbonyl group. These values are synonymous with previously reported coumarin ligands (1.351, 1.432 and 1.432 Å respectively).[42] The methylene group accommodates in the same plane as the alkynyl moieties with angles range between 175.9(7)º and 178.4(10)°, which are similar to the values found in literature (177.6(2)º).[35]

The Au−C and C−C distances in the alkyne unit are observed to be, for the former, in the lower values that in previously reported alkyne derivatives and the latter similar than in the similar gold(I) alkynyl complexes.[35,43-46] As shown in Figure 6, the ligand molecules coordinated to linear geometry Au(I) are twisted in orthogonal fashion to avoid steric hindrance with torsional angles 111.30(7)° [P1-Au2-Au1-P2] and 120.20(7)° [P3-Au3-Au2-P1]. Subsequently, the linearly coordinated Au(I) ions self-assembles *via* Au···Au interactions at distances f 3.1274(4) Å [Au1-Au2] and 3.1316(4) Å [Au2-Au3] to give a zigzag 1-D polymeric chain, as shown in Figure 7. The aurophilic interactions observed in complex **1** are similar to those observed in gold(I) derivatives with PTA[43,44,47,48] and shorter than values found similar 1-D polymeric complexes of $[Au(PTA)_4][Au(CN_2)]$[49] and $[Au(C{\equiv}C\text{-}CH_2\text{-}SC_5H_4N)(PTA)]$[50] with distances *ca.* 3.45 Å and *ca.* 3.27 Å, respectively.

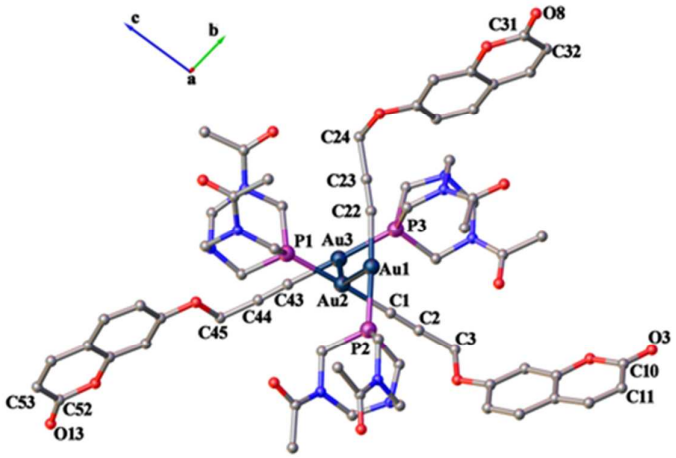

Figure 6. Asymmetric unit of complex **1** viewed down the *a*-axis displaying aurophilic interactions. Hydrogen atoms and solvent molecules are omitted for clarity.

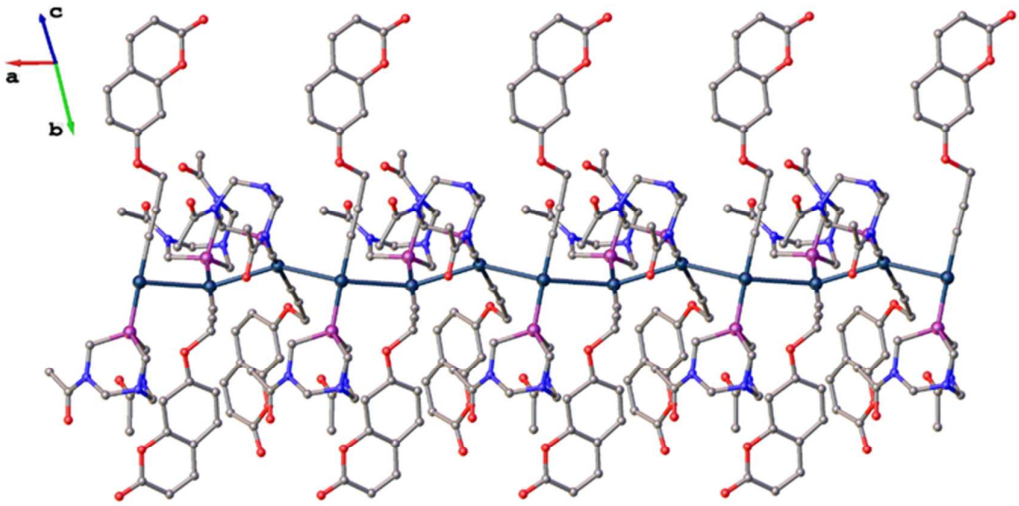

Figure 7. 1-D Polymeric structure of complex **1**, hydrogen atoms and solvent molecules are omitted for clarity.

In the 3-D crystal packing of complex 1, the planar coumarin ligands in each 1-D polymeric chain are positioned in three orientations to give star-shape (Figures S4 and

S5) motif around the irregular aurophilic wire exhibiting π-π interactions, with centroid-centroid distances of (Figure S3) 3.69(3) Å. These distances are similar to our anionic gold(I) dialkynyl-coumarin complex (3.646 Å) very recently reported by us.[35]

H-bondings are also observed between the methylene groups and the acetyl units of the DAPTA phosphine with O···H distance ca. 2.34 - 2.79 Å.

All these weak interactions (aurophilic, π–π stacking and hydrogen bondings) contributes complex **1** to pack three-dimensionally to give MOF-like structure, as shown in Figure 8. Curiously, to the best of our knowledge, porous network with 1-D aurophilic interactions that self-assembled either linearly or non-linearly constituting aromatic ligands, and stabilized through such interactions to generate MOF-like structure has no precedent in the literature.

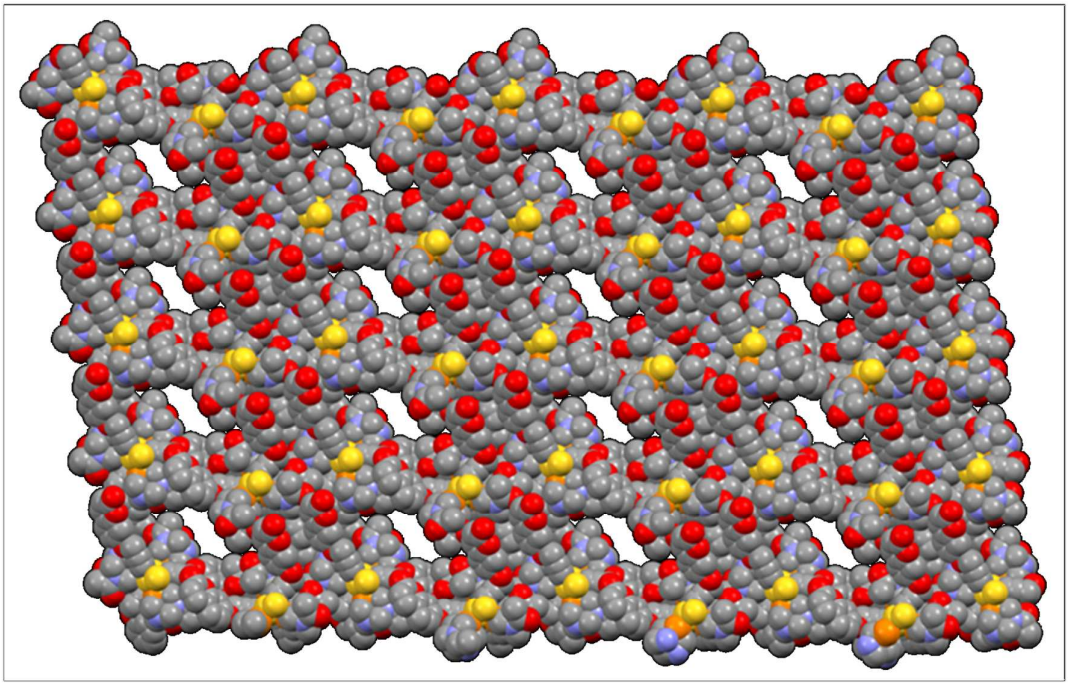

Figure 8. Space fill representation of the 3-D crystal packing of complex **1** viewed down the *a*-axis displaying channels for MOF-like motif. Hydrogen atoms and dichloromethane molecules are omitted for clarity.

Taking into consideration this data, we could expect hydrogen-bonding and π -π stacking interactions to be more favoured in water, being the driving force of the formation of gels. On the other hand, theoretical calculations carried out with similar aurophilic hydrogelators indicate a clear involvement of Au(I)···Au(I) interactions in the formation of these large aggregates.[34]

Conclusions

The design and characterization of a alkynyl gold(I) complex with a coumarin fluorophore as a new hydrogelator for the synthesis of luminescent soft materials has been performed.

Absorption and DLS studies in aqueous media suggested the formation of aggregates at low micromolar concentrations. The same techniques indicated that aggregation was favoured with increasing temperature.

Fluorescence microscopy revealed the presence of luminescent fiber-like structures at milimolar concentrations. While strong blue emission (corresponding to the fluorescence of the coumarin) was only observed in isolated fibers, green emission (derived from coumarin phosphorescence) was observed throughout the bulk of the reticulated network of fibers. This effect is accounted by the spatial proximity of coumarins and gold(I) atoms which leads an increase in Inter-System Crossing and consequently, quenching of fluorescence and enhancement of triplet emission ($T_1 - S_0$ radiative pathway).

X-Ray crystallography confirmed the existence of π-π interactions, hydrogen bondings as well as aurophilic interactions, resulting in highly ordered linear assemblies.

Acknowledgements

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Supporting Information Available

Hydrogel formation with complex **1** checked by the inverted tube method (Figure S1); Absorption spectra of **1** between 5 to 100 μM in water. Inset: Plot of the absorption at 400 nm *vs*. concentration (Figure S2); DLS auto-correlation data (left) and resulting size distributions (right) of aqueous solutions of **1** at different concentrations in water (Figure S3); Schematic representation of (a) star-shaped 1-D polymeric structure of **1** and (b) observed π–π interactions (Figure S4); 3-D crystal packing of **1** (Figure S5). X-ray crystallographic data for **1** (Table S1).

References

[1] S. Dong, B. Zheng, D. Xu, X. Yan, M. Zhang and F. Huang, *Adv. Mater*. 2012, **24**, 3191.

[2] M. Fleischer and C. Schmuck, *Chem. Commun.* 2014, **50**, 10464.

[3] S. Banerjee, R. K. Das and U. Maitra, *J. Mater. Chem.* 2009, **19**, 6649.

[4] A. R. Hirst, B. Escuder, J. F. Miravet and D. K. Smith, *Angew. Chem. Int. Ed.* 2008, **47**, 8002.

[5] N. M. Sangeetha and U. Maitra, *Chem. Soc. Rev.* 2005, **34**, 821.

[6] T. Vermonden, R. Censi and W. E. Hennink, *Chem. Rev*., 2012, **112**, 2853

[7] E. A. Appel, J. del Barrio, X. J. Loh and O. A. Scherman, *Chem. Soc. Rev.*, 2012, **41**, 6195.

[8] D.D. Díaz, D. Ku¨hbeck and R. J. Koopmans, *Chem. Soc. Rev*., 2011, **40**, 427.

[9] S. Saha, J. Bachl, T. Kundu, D. Díaz Díaz and R. Banerjee, *Chem. Commun.,* 2014, **50**, 7032.

[10] S.-T. Lam, G. Wang and V.W.-W. Yam, *Organometallics* 2008, **27**, 4545.

[11] M.-O. M. Piepenbrock, G. O. Lloyd, N. Clarke and J. W. Steed, *Chem. Rev*., 2010, **110**, 1960.

[12] A. Yiu-Tam and V. W.-W. Yam, *Chem. Soc. Rev*., 2013, **42**, 1540.

[13] J. Zhang and C.-Y. Su, *Coord. Chem. Rev.*, 2013, **257**, 1373.

[14] J. H. Jung, J. H. Lee, J. R. Silverman and G. John, *Chem. Soc. Rev*., 2013, **42**, 924.

[15] T.H.T. Hsu, J. J. Naidu, B.-J. Yang, M.-Y. Jang, I. J. B. Lin, *Inorg. Chem*. 2012, **51**, 98.

[16] V. K.-M. Au, N. Zhu and V. W.-W. Yam, *Inorg. Chem.* 2013, **52**, 558.

[17] R. Gavara, J. Llorca, J.C. Lima and L. Rodríguez, *Chem. Comm.* 2013, **49**, 72.

[18] E. Aguiló, R. Gavara, J.C. Lima, J. Llorca and L. Rodríguez, *J. Mat. Chem. C* 2013, **1**, 5538.

[19] T.F.A. de Greef, M.M.J. Smudlers, M. Wolffs, A.P.H.J. Schenning, R.P. Sijbesma and E.W. Meijer, *Chem. Rev*. 2009, **109**,5687.

[20] S. Sculfort and P. Braunstein, *Chem. Soc. Rev.* 2011, **40**, 2741.

[21] H. Schmidbaur and A. Schier, *Chem. Soc. Rev.* 2012, **41**, 370.

[22] J.C. Lima and L. Rodríguez, *Chem. Soc. Rev*. 2011, **40**, 5442.

[23] R. Visbal and M. C. Gimeno, *Chem. Soc. Rev.,* 2014, **43**, 3551.

[24] R. Visbal, I. Ospino, J. M. L ópez-de-Luzuriaga, A. Laguna, and M.C. Gimeno, *J. Am. Chem. Soc*., 2013, **135**, 4712.

[25] E.R.T. Tiekink and J.-G. Kang, *Coord. Chem. Rev*. 2009, **253**, 1627.

[26] X. He and V.W.-W. Vivian, *Coord. Chem. Rev*. 2011, **255**, 2111.

[27] L. Rodríguez, J.C. Lima, M. Ferrer, O. Rossell, M. Engeser, *Inorg. Chimica Acta* 2012, **381**, 195.

[28] R. J. Puddephatt, *Chem. Soc. Rev*., 2008, **37**, 2012.

[29] M.J. Katz, K. Sakai and D.B. Leznoff, *Chem. Soc. Rev*. 2008, **37**, 1884.

[30] H.E. Abdou, A.A. Mohamed, J.P. Fackler, A. Burini, R. Galassi, J.M. Lopez-de-Luzuriaga and M.E. Olmos, *Coord. Chem. Rev*. 2009, **253**, 1661.

[31] M. Ferrer, A. Gutiérrez, M. Mounir, L. Rodríguez, O. Rossell, M. Font-Bardia, P. Gómez-Sal, A. Martín and X. Solans, *Organometallics* 2011, **30**, 3419.

[32] E.R.T. Tiekink, *Coord. Chem. Rev.* 2014, **275**, 130.

[33] H. Schmidbaur and A. Schier, *Gold Bull*. 2000, **33**, 3.

[34] R. Gavara, E. Aguiló, D. Kubániová, C. Fonseca Guerra, L. Rodríguez and J.C. Lima, *manuscript in preparation.*

[35] J.Arcau, V.Andermark, E.Aguiló, A.Gandioso, A.Moro, M. Cetina, J.C. Lima, K.Rissanen, I. Ott and L. Rodríguez, *Dalton Trans.* 2014, **43**, 4426.

[36] Oxford Diffraction, Xcalibur CCD System. *CRYSALISPRO*; Oxford Diffraction Ltd: Abingdon, England, 2013.

[37] M. C. Burla, M. Camalli, B. Carrozzini, G. L. Cascarano, C. Giacovazzo, G. Polidori and R. Spagna, *J. Appl. Crystallogr.* 2003, **36**, 1103.

[38] G. M. Sheldrick, *Acta Crystallogr.* 2008, **A64**, 112.

[39] L. J. Farrugia, *J. Appl. Crystallogr.* 1999, **32**, 837.

[40] N. Jordão, R. Gavara, A.J. Parola, *Macromolecules* 2013, **46**, 9055.

[41] E. García-Moreno, S. Gascón, M.J. Rodriguez-Yoldi, E. Cerrada and M. Laguna, *Organometallics* 2013, **32**, 3710.

[42] K. Ueno, *Acta Cryst.* 1985, **41**, 1786.

[43] E. Vergara, E. Cerrada, A. Casini, O. Zava, M. Laguna and P.J. Dyson, *Organometallics* 2010, **29**, 2596.

[44] F. Mohr, E. Cerrada and M. Laguna, *Organometallics* 2006, **25**, 644.

[45] M. Ferrer, A. Gutiérrez, L. Rodríguez, O. Rossell, J.C. Lima, M. Font-Bardia and X. Solans, *Eur. J. Inorg. Chem*. 2008, 2899.

[46] M. Ferrer, L. Rodríguez, O. Rossell, F. Pina, J.C. Lima, M. Font-Bardia and X. Solans, *J. Organomet. Chem.* 2003, **678**, 82.

[47] J.M. Forward, D. Bohmann, J.P. Fackler and R.J. Staples, *Inorg. Chem.* 1995, **34**, 6330.

[48] Z. Assefa, B.G. McBurnett, R.J. Staples and J.P. Jr. Fackler, *Inorg. Chem.* 1995, **34**, 4965.

[49] Z. Assefa, M. A. Omary, B. G. McBurnett, A. A. Mohamed, H. H. Patterson and R. J. Staples, J. P. Fackler, *Inorg. Chem*., 2002, **41**, 6274.

[50] E. García-Moreno, S. Gascón, M. J. Rodríguez-Yoldi, E. Cerrada and M. Laguna, Organometallics, 2013, **32**, 3710.