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Complete List of Authors:	Perkin, Susan; University of Oxford, Department of Chemistry Cousens, Nico; Oxford University, Chemistry Kearney, Leah; Oxford University, Chemistry Clough, Matthew; Imperial College, London Lovelock, Kevin; Imperial College London, Chemistry Palgrave, Robert; UCL, Department of Chemistry

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Nico E. A. Cousens^a, Leah J. Taylor Kearney^{a†}, Matthew T. Clough^b, Kevin R. J. Lovelock^b, Robert G. Palgrave^c, Susan Perkin^{*a}

Graphical Abstract



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Nico E. A. Cousens^a, Leah J. Taylor Kearney^{a†}, Matthew T. Clough^b, Kevin R. J. Lovelock^b, Robert G. Palgrave^c, Susan Perkin^{*a}

^a*Department of Chemistry, Physical and Theoretical Chemistry Laboratory, University of Oxford, South Parks Road, Oxford OX1 3QZ, UK. E-mail: susan.perkin@chem.ox.ac.uk*

^b*Department of Chemistry, Imperial College London, London SW7 2AZ, UK*

^c*Department of Chemistry, University College London, London WC1H 0AJ, UK*

[†]Current address: School of Chemistry, National University of Ireland, Galway, University Road, Galway, Ireland

Abstract

A range of ionic liquids containing dialkylimidazolium cations and halobismuthate anions ($[\text{BiBr}_x\text{Cl}_y\text{I}_z]^-$ and $[\text{Bi}_2\text{Br}_x\text{Cl}_y\text{I}_z]^-$) were synthesised by combining dialkylimidazolium halide ionic liquids with bismuth (III) halide salts. The majority were room temperature liquids, all with very high densities. The neat ionic liquids and their mixtures with 1-butyl-3-methylimidazolium *bis*(trifluoromethanesulfonyl)imide were characterised using Densitometry, Viscometry, NMR Spectroscopy, Electrospray Ionisation Mass Spectrometry (ESI), Liquid Secondary Ion Mass Spectrometry (LSIMS), Matrix-assisted Laser Desorption/Ionization Mass Spectrometry (MALDI), X-Ray Photoelectron Spectroscopy (XPS) and Thermogravimetric Analysis (TGA), to establish their speciation and suitability for high-temperature applications.



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1 Introduction

Anions consisting of a metal centre coordinated by halide ions are well known in the field of ionic liquids. Indeed, the tetrachloroaluminate anion, $[\text{AlCl}_4]^-$, was discovered in the late 1950s and was

the first room temperature 'molten salt' anion.^[1] Many similar ions have since been reported such as tetrachloroferrate, $[\text{FeCl}_4]^-$; tetrabromoferrate, $[\text{FeBr}_4]^-$; and tetrachlorogallate, $[\text{GaCl}_4]^-$.^[2] When paired with suitable cations, these form what are commonly termed the 'halometallate' family of ionic liquids. Traditionally seen as solvents for Lewis acidic catalysis and electrochemistry, their applications are broadening, and they now find use in areas as diverse as biomass processing and gas storage.^[3] This paper reports bismuth based analogues, which have the form $[\text{BiX}_4]^-$, where X is Cl^- , Br^- , I^- or a mixture of these ions. These bismuthate anions were combined with 1-ethyl-3-methylimidazolium ($[\text{C}_2\text{C}_1\text{Im}]^+$) and 1-butyl-3-methylimidazolium ($[\text{C}_4\text{C}_1\text{Im}]^+$) cations, which are known to readily form room temperature liquids.^[4] Mixtures with the ionic liquid 1-butyl-3-methylimidazolium *bis*(trifluoromethylsulfonyl)imide ($[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$) were also investigated. The chemistry of bismuth in chloroaluminates has been investigated by Ahmed, Köhler and Ruck, however, to the knowledge of the authors there have been no previous reports of pure halobismuthate ionic liquids.^[5]

The most common halometallate ionic liquids consist of a group 13 metal centre. Group 13 metal halides are good Lewis acids and as such readily form complexes with electron donating halides, as is the case for $[\text{AlCl}_4]^-$. Bismuth also prefers to adopt a 3+ oxidation state due to the inert pair effect and is similarly Lewis acidic due to weak shielding of its nucleus by its 4f electrons.^[6] This makes it an excellent candidate as a possible halometallate ion centre, prepared by the addition of a halide ion to a bismuth halide salt. Due to the high molecular weight of bismuth, it was anticipated that the resulting liquids would have higher densities than other metal based ionic liquids.

The halometallates can form various different anionic species in solution depending on, for instance, the amount of AlCl_3 added to a chloroaluminate; Cl^- , $[\text{AlCl}_4]^-$, $[\text{Al}_2\text{Cl}_7]^-$ and $[\text{Al}_3\text{Cl}_{10}]^-$ species are all known to exist.^[7] Given the chemical similarity of bismuth halides to the group 3 halides, the speciation in halobismuthate ionic liquids may be equally interesting and was investigated using Electrospray Ionisation Mass Spectrometry (ESI), Liquid Secondary Ion Mass Spectrometry (LSIMS), Matrix-assisted Laser Desorption/Ionization Mass Spectrometry (MALDI) and X-Ray Photoelectron Spectroscopy (XPS).

A common issue with ionic liquids is that they can be very viscous, which was found to be the case for the halobismuthate ionic liquids investigated in this study. Ionic liquids incorporating the *bis*(trifluoromethanesulfonyl)imide, $[\text{NTf}_2]^-$, anion are favoured for their low viscosities.^[8] Therefore, binary ionic liquid mixtures of one particular halobismuthate ionic liquid, 1-butyl-3-methylimidazolium tribromochlorobismuthate, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, were prepared by mixing $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$, in an attempt to reduce the viscosity yet retain any attractive chemical and physical properties of the new liquids. An alternative solution is to work with the ionic liquids at elevated temperatures. Therefore, the thermal stability of the liquids was studied by Thermogravimetric Analysis and their viscosities as a function of temperature are also reported.

2 Results and discussion

2.1 Synthesis and stability

A range of ionic liquids incorporating the $[C_2C_1Im]^+$ and $[C_4C_1Im]^+$ cations and $[BiBr_xCl_yI_z]^-$ or $[Bi_2Br_xCl_yI_z]^-$ anions, where x , y and z indicate varying stoichiometries of each halide, were successfully synthesised (Figure 1 and Table 1). At 100 °C all were translucent, homogenous liquids. Chloride-containing liquids were colourless, bromide-containing liquids yellow and iodide-containing liquids red/purple in appearance. Upon cooling to room temperature certain liquids, as reported in Table 1, became opaque and solidified. Of the wide range of bismuthate ionic liquids prepared, $[C_4C_1Im][BiBr_3Cl]$, $[C_4C_1Im][BiBr_3I]$, $[C_4C_1Im][Bi_2Br_7]$ and $[C_4C_1Im][Bi_2BrI_6]$, as well as mixtures of $[C_4C_1Im][BiBr_3Cl]$ with $[C_4C_1Im][NTf_2]$, were chosen for further characterisation as these ionic liquids cover a broad variety of anions.

$[C_4C_1Im][BiBr_3Cl]$

1H NMR (400 MHz, Acetone- d_6): δ 9.13 (1H, s), 7.75 (1H, t, $J = 2$ Hz), 7.71 (1H, t, $J = 2$ Hz), 4.43 (2H, t, $J = 7$ Hz), 4.11 (3H, s), 2.01 - 1.90 (2H, m), 1.46 - 1.34 (2H, m), 0.95 (3H, t, $J = 7$ Hz). ^{13}C NMR (100 MHz, Acetone- d_6): δ 137.2, 124.9, 123.4, 50.5, 37.4, 32.9, 20.1, 13.8. m/z (ESI $^+$): 139 (100%) $[C_4C_1Im]^+$, 83 (7%) $[C_1HIm]^+$. m/z (ESI $^-$): 515 (100%) $[BiBr_4]^-$, 298 (69%) $[C_4C_1ImBr_2]^-$, 471 (36%) $[BiBr_3Cl]^-$, 372 (13%) $[BiBr_2]^-$. m/z (LSIMS $^+$): 139 (100%) $[C_4C_1Im]^+$, 83 (51%) $[C_1HIm]^+$. m/z (LSIMS $^-$): 79 (100%) $^{79}Br^-$, 81 (97%) $^{81}Br^-$, 489 (2%) $[BiBr_3Cl]^-$. m/z (MALDI $^-$): 528 (100%) $[BiBr_4]^-$, 484 (71%) $[BiBr_3Cl]^-$, 294 (39%) $[C_4C_1ImBr_2]^-$, 438 (26%) $[BiBr_2Cl_2]^-$, 637 (2%) $[C_4C_1ImBi_2Br]^-$.

$[C_4C_1Im][BiBr_3I]$

1H NMR (400 MHz, Acetone- d_6): δ 9.11 (1H, s), 7.76 (1H, t, $J = 2$ Hz), 7.72 (1H, t, $J = 2$ Hz), 4.43 (2H, t, $J = 8$ Hz), 4.12 (3H, s), 2.02 - 1.91 (2H, m), 1.47 - 1.35 (2H, m), 0.95 (3H, t, $J = 7$ Hz). ^{13}C NMR (100 MHz, Acetone- d_6): δ 137.1, 124.9, 123.5, 50.6, 37.4, 32.9, 20.1, 13.8. m/z (ESI $^+$): 139 (100%) $[C_4C_1Im]^+$, 83 (2%) $[C_1HIm]^+$. m/z (ESI $^-$): 517 (100%) $[BiBr_4]^-$, 563 (70%) $[BiBr_3I]^-$, 298 (47%) $[C_4C_1ImBr_2]^-$, 609 (19%) $[BiBr_2I_2]^-$. m/z (LSIMS $^+$): 83 (74%) $[C_1HIm]^+$, 139 (100%) $[C_4C_1Im]^+$. m/z (LSIMS $^-$): 79 (100%) $^{79}Br^-$, 81 (97%) $^{81}Br^-$, 127 (31%) I^- , 531 (12%) $[BiBr_4]^-$, 580 (10%) $[BiBr_3I]^-$, 452 (7%) $[BiBr_3]^-$, 371 (6%) $[BiBr_2]^-$, 498 (6%) $[BiBr_2I_2]^-$, 160 (5%) $[Br_2]^-$, 207 (2%) $[BrI]^-$. m/z (MALDI $^-$): 528 (100%) $[BiBr_4]^-$, 574 (69%) $[BiBr_3I]^-$, 637 (99%) $[C_4C_1ImBi_2Br]^-$.

$[C_4C_1Im][Bi_2Br_7]$

m/z (LSIMS $^+$): 83 (72%) $[C_1HIm]^+$, 139 (73%) $[C_4C_1Im]^+$. m/z (LSIMS $^-$): 527 (100%) $[BiBr_4]^-$, 446 (32%) $[BiBr_3]^-$, 367 (23%) $[BiBr_2]^-$, 293 (8%) $[C_4C_1ImBiBr_3]^{2-}$, 657 (5%) $[Bi_2Br_3]^-$, 1197 (5%) $[Bi_3Br_7]^-$, 740 (3%) $[Bi_2Br_4]^-$, 815 (3%) $[Bi_2Br_5]^-$, 976 (2%) $[Bi_2Br_7]^-$.

$[C_4C_1Im][Bi_2BrI_6]$

m/z (LSIMS $^+$): 83 (53%) $[C_1HIm]^+$, 139 (67%) $[C_4C_1Im]^+$. m/z (LSIMS $^-$): 293 (100%) $[C_4C_1ImBiBr_3]^{2-}$, 527 (63%) $[BiBr_4]^-$, (62%) 714 $[BiI_4]^-$, 339 (48%) $[BiI]^-$, 588 (48%) $[BiI_3]^-$, 461 (37%) $[BiI_2]^-$, 669 (28%) $[BiBrI_3]^-$, 413 (15%) $[BiBrI]^-$, 446 (27%) $[BiBr_3]^-$, 620 (13%) $[Bi_2BrI]^-$, 494 (8%) $[BiBr_2I]^-$, 1051 (4%) $[Bi_2I_5]^-$, 797 (4%) $[Bi_2I_3]^-$, 749 (3%) $[Bi_2BrI_2]^-$, 830 (3%) $[Bi_2Br_2I_2]^-$, 1487 (3%) $[Bi_3Br_6I_3]^-$.

1H and ^{13}C NMR spectra and ESI, LSIMS and MALDI mass spectra confirm the presence of the $[C_4C_1Im]^+$ cation. Speciation of the anions is discussed in section 2.7. It should be noted that many of

the peaks in the mass spectra could not be assigned (see supplementary information, Figures S1 to S14), and those that were, were done so with caution. ^{35}Cl and ^{79}Br NMR experiments were attempted to see if it would be possible to distinguish different environments for the halide ligands in axial and equatorial positions, however, no peaks were distinguishable from the baseline noise.

When mixed with water the liquids formed a white solid, presumably containing mixed bismuth (III) oxy halide compounds, however no degradation of the liquids was observed when left open to the air even after a period of several months. This is in stark contrast to the well-understood haloaluminate ionic liquids, which hydrolyse immediately upon contact with moisture in the air. It should also be noted that bismuth itself, despite being a 'heavy metal', is non-toxic.^[9]

2.2 Density

The densities of the liquids are remarkably high (Table 1), presumably due in part to the high molar mass of bismuth (bismuth is the heaviest stable element in the periodic table). The densest liquid in the series, $[\text{C}_4\text{C}_1\text{Im}][\text{Bi}_2\text{BrI}_6]$, has a density of $3.46 \pm 0.05 \text{ g cm}^{-3}$ (at $100 \text{ }^\circ\text{C}$), although this was not a liquid at room temperature. This density is more than double that of common ionic liquids (*e.g.* $[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2] = 1.43 \text{ g cm}^{-3}$).^[8] According to Welton and Wasserscheid, the highest known ionic liquid density is 2.4 g cm^{-3} for $[(\text{CH}_3)_3\text{S}][\text{Al}_2\text{Br}_7]$.^[8] The densest room temperature liquid in our investigation was $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}][\text{BiBr}_3]_{0.5}$, with a density of $2.65 \pm 0.05 \text{ g cm}^{-3}$.

The density of the liquids correlates with the molecular weight of the halide ligands: $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ had a density of $2.41 \pm 0.05 \text{ g cm}^{-3}$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$ had a density of $2.49 \pm 0.05 \text{ g cm}^{-3}$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ had a density of $2.63 \pm 0.05 \text{ g cm}^{-3}$. This trend is also seen in the literature for haloaluminates ($[\text{C}_2\text{C}_1\text{Im}][\text{AlCl}_4] = 1.27 \text{ g cm}^{-3}$, $[\text{C}_2\text{C}_1\text{Im}][\text{AlBr}_4] = 1.92 \text{ g cm}^{-3}$ at $50 \text{ }^\circ\text{C}$) and haloferrates ($[\text{C}_4\text{C}_1\text{Im}][\text{FeCl}_4] = 1.38 \text{ g cm}^{-3}$, $[\text{C}_4\text{C}_1\text{Im}][\text{FeBr}_4] = 1.98 \text{ g cm}^{-3}$ at $25 \text{ }^\circ\text{C}$).^[10]

2.3 Thermal stability

Halobismuthate ionic liquids may find use in high-temperature applications, and as such it is important to understand their thermal stability. The thermal stabilities of $[\text{C}_4\text{C}_1\text{Im}][\text{BiCl}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]_{0.5}[\text{NTf}_2]_{0.5}$ were measured using temperature-ramped Thermogravimetric Analysis (TGA), and are compared in Figure 2 and Table 2. The T_{onset} thermal decomposition temperatures were determined from extrapolation of the steepest point of the thermograph. T_{onset} temperatures of the measured halobismuthate liquids, with the exception of $[\text{C}_4\text{C}_1\text{Im}][\text{BiCl}_4]$, are similar and high (between $370 \text{ }^\circ\text{C}$ to $380 \text{ }^\circ\text{C}$). By contrast, $[\text{C}_4\text{C}_1\text{Im}][\text{BiCl}_4]$ decomposed in a multi-step fashion, with a T_{onset} value of $260 \text{ }^\circ\text{C}$ corresponding to the first step of weight loss. This T_{onset} value compares reasonably favourably with that of the ionic liquid $[\text{C}_4\text{C}_1\text{Im}]\text{Cl}$ (measured as $275 \text{ }^\circ\text{C}$ under the same experimental conditions), suggesting that the chloride ligands in the $[\text{BiCl}_4]^-$ anion may be more labile and free to attack the dialkylimidazolium cation than in the other halobismuthate analogues.

The addition of $[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$ to reduce the viscosity (section 2.5) appears to have no detrimental effect on the thermal stability. This implies that these mixtures are suitable for use at elevated temperatures. It is well established that T_{onset} dramatically overestimates the thermal stability of ionic liquids, but is still a useful parameter for comparing the relative stabilities of ionic liquids.^[11]

However, the very high T_{onset} values (> 370 °C) for the halobismuthate ionic liquids suggest that they would be stable at temperatures in excess of 200 °C for extended periods of time.

It is unclear whether TGA mass loss is due to decomposition or vaporisation of the liquids, or a combination of these processes. The fact that T_{onset} values are well below the boiling points of the composite bismuth halides (BiCl_3 , BiBr_3 or BiI_3 which boil at 447 °C, 453 °C and 542 °C respectively) suggests that mass loss is primarily due to decomposition.^[12]

2.4 Glass transition temperature (T_g)

The liquids $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}][\text{BiBr}_3]_{0.5}$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ are room temperature liquids (Table 1). Upon decreasing the length of the alkyl chain on the imidazolium ring from $[\text{C}_4\text{C}_1\text{Im}]^+$ to $[\text{C}_2\text{C}_1\text{Im}]^+$ with the $[\text{BiBr}_3\text{Cl}]^-$ anion, T_g increases, possibly due to closer packing of the ions and hence stronger Coulomb attractions. $[\text{C}_4\text{C}_1\text{Im}][\text{Bi}_2\text{Br}_6\text{Cl}]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ are liquids above 50 °C while $[\text{C}_4\text{C}_1\text{Im}][\text{Bi}_2\text{Br}_6\text{I}]$ is a liquid above 100 °C. This trend of increasing T_g with increasing I^- content can be explained by the increasing van der Waals attractions between the more polarisable I^- ligands.^[13] The T_g of all the ionic liquids was well below the melting points of the component compounds, suggesting that new chemical species had formed.

2.5 Viscosity

The room temperature viscosities of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ were measured (Table 1). The viscosity of iodide-containing $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ is greater than its chloride analogue, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ (1517 Pa s and 648 Pa s at 25 °C, respectively), although both liquids are still very viscous. As described above, the liquid $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ offers the following interesting and beneficial properties: (i) very high density; (ii) high thermal stability; (iii) a melting point below room temperature; (iv) ease of preparation. However, the very high viscosity of this ionic liquid limits its use at room temperature. Therefore, the effect of adding the prototypical ionic liquid $[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$, which has a much lower viscosity of 0.052 Pa s, was investigated.^[8] The density of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]/[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$ mixtures is shown in Figure 3a. A 50:50 mol% ratio lowers the viscosity of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ from 648 Pa s to 1.37 Pa s, making it convenient to handle at room temperature. Thus, the viscosity of these mixtures was far lower than the parent halobismuthate ionic liquid, yet the density was still high and the thermal stability was not compromised (Figure 2). Below 50 mol% of the halobismuthate ionic liquid, the mixtures were no longer stable and the $[\text{BiBr}_3\text{Cl}]^-$ ion appeared to separate into Cl^- and BiBr_3 , which precipitated as a yellow solid.

Alternatively, the viscosity can be reduced by increasing the temperature. The effect of temperature on viscosity is shown in Figure 3b. As expected, there is a dramatic decrease in viscosity as the temperature is increased. At 85 °C the viscosity of pure $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ drops to 2.82 Pa s. Given the good thermal stability of the liquids, raising the temperature is clearly an effective method for reducing the viscosity of halobismuthate liquids.

As water content is known to have a substantial effect on viscosity, prior to viscosity measurements the liquids were dried overnight in a desiccator at <0.1 mbar and then stored under dry nitrogen. The water content of the liquids was measured immediately after exposure to the ambient atmosphere and then again after 10 minutes and 30 minutes.^[8] Water uptake was extremely slow; the water content varied from 224 ppm to 470 ppm for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]_{0.5}[\text{NTf}_2]_{0.5}$.

2.6 X-ray Photoelectron Spectroscopy (XPS)

Two pieces of information available from XPS are of use in this study: composition and electron binding energies. Binding energies are used to determine the number of electronic environments present in the samples, and are of primary interest here. In addition, the composition by XPS is compared to the expected bulk stoichiometric composition to check that the samples were of high purity.

For the composition, relative sensitivity factors (RSFs) from Kolbeck *et al.* are used for all elements apart from bismuth (Table 3). These RSFs were determined for a range of ionic liquids.^[14] Note that the Kolbeck RSFs for Cl and Br were determined for Cl 2p and Br 3d peak areas, whereas here the Cl 2p_{3/2} and Br 3d_{5/2} peak areas are used; therefore, the Kolbeck RSFs are multiplied by 2/3 to account for the different peaks used. The RSF for Bi used is 4.25, taken from Wagner *et al.*^[15] The error in experimental composition is estimated as ±10%. The information depth (ID) of XPS experiments can be defined as the depth, within the sample, from which 95% of the measured signal will originate. ID is assumed to vary mainly with $\cos\theta$, where θ is the electron emission angle relative to the surface normal. If it is assumed that the inelastic mean free path (λ) of photoelectrons in organic compounds is of the order of ~3 nm, at the kinetic energies employed here the ID in this geometry can be estimated, when $\theta = 0^\circ$, ID = 7–9 nm.^[16] Consequently, these data may be considered as representative of the bulk composition and do not reflect any local enhancements of concentration at the near surface region.

When comparing the nominal and XPS experimental compositions, it is best to start with a reference ionic liquid, [C₈C₁Im]Br. The nominal and XPS experimental compositions are the same, within the error of the experiment, for the carbon and nitrogen regions. The nominal and XPS experimental compositions are slightly different for Br for [C₈C₁Im]Br, indicating that the Br 3d_{5/2} RSF is slightly too large. If RSF(Br 3d_{5/2}) = 0.330 is used, the Br nominal and XPS experimental compositions match very well. Overall, the good match between nominal and XPS experimental compositions confirms the reliability of the RSFs used, the purity of [C₈C₁Im]Br, and the XPS system used here. For the four halobismuthate ionic liquids, the nominal and XPS experimental compositions match well. As for [C₈C₁Im]Br, the experimental XPS number of atoms for Br is in all cases slightly lower than the nominal number of atoms. Again, if RSF(Br 3d_{5/2}) = 0.330 is used, the Br nominal and XPS experimental compositions match very well for the halobismuthate ionic liquids. The experimental XPS numbers of atoms for bismuth is in all cases slightly lower than the nominal number of atoms. This observation could be due to there being less bismuth in the ionic liquids than expected, or because the RSF(Bi 4f_{7/2}) is slightly too low. As the result is consistent across all four halobismuthate ionic liquids, it can be concluded that the RSF(Bi 4f_{7/2}) is slightly too low. It should be noted that there is a small amount of bismuth metal contamination (ratio of peak areas for Bi(III):Bi(metal) is ~6.1) for [C₄C₁Im][BiBr₄]_{0.67}Br_{0.33}. This metal contamination is very likely due to sample degradation by the X-ray source, not due to the presence of bismuth metal in the sample at the start of the experiment. Such sample degradation is commonly observed for XPS.^[17] The bismuth region was recorded at the end of the XPS experiment; therefore, the results for [C₄C₁Im][BiBr₄]_{0.67}Br_{0.33} can be used to probe the electronic environment of [C₄C₁Im][BiBr₄]_{0.67}Br_{0.33} with confidence. Due to the relatively small amount of bismuth metal contamination, it is not expected that this metal had any effect on the XPS results for the other regions for [C₄C₁Im][BiBr₄]_{0.67}Br_{0.33}. All elements expected to be present in the samples are observed, and no extra elements are observed, aside from a small

amount of oxygen contamination for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]_{0.67}\text{Br}_{0.33}$ (~0.5 atoms in the sample). This oxygen contamination is likely to be due to X-ray degradation of the sample, or the presence of water impurity in the solid. Therefore, it can be concluded that the compositions of the ionic liquids probed by XPS match well with the nominal compositions; therefore, information about the electronic environments of the ionic liquids can be obtained with good confidence that ionic liquids were of high purity.

To be able to directly compare binding energies for different ionic liquids, it is necessary to charge correct all XP spectra. For a wide range of imidazolium-based ionic liquids, the C_{alkyl} 1s peak has successfully been used for charge correction.^[18] For all five ionic liquids studied here, two distinct peaks with a shoulder at higher energies are observed for C 1s region (Figures S15-S19), as expected based upon previous results for XPS of imidazolium-based ionic liquids.^[18] In addition, there is a very broad shake-up feature at ~293 eV, also as expected for $[\text{C}_n\text{C}_1\text{Im}]^+$ -based ionic liquids (where $n \geq 2$).^[18c] Therefore, there are three different electronic environments for carbon atoms in all five of the ionic liquids studied here. The lower energy peak is assigned as C_{alkyl} 1s; the higher energy peak and shoulder are assigned as C_{hetero} 1s. All C 1s XP spectra are fitted with three peaks, and give a very good fit (Figures S15-S19). The binding energy of the C_{alkyl} 1s peak is set to 285.0 eV, and all of the other peaks are corrected by the same amount. This charge correction procedure allows comparison of the binding energies for different ionic liquids. Using C_{alkyl} 1s for charge correction for $[\text{C}_4\text{C}_1\text{Im}][\text{A}]$ gives an error of ± 0.2 eV, larger than the ± 0.1 eV error associated with studying $[\text{C}_8\text{C}_1\text{Im}][\text{A}]$. Binding energies from the XP spectra of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]_{0.67}\text{Br}_{0.33}$ and $[\text{C}_8\text{C}_1\text{Im}]\text{Br}$ are shown in Table 4.

The three ionic liquids that are viscous liquids at room temperature ($[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$) give two peaks in the Bi 4f region, $4f_{5/2}$ and $4f_{7/2}$, which are due to spin orbit coupling (Figure 4a). Therefore, for these three ionic liquids there is one bismuth electronic environment only. The binding energies for the Bi $4f_{7/2}$ peak for all three ionic liquids are the same (within the error of the experiment), 159.8 ± 0.1 eV. These binding energies are very similar to neutral reference bismuth(III) halide compounds BiF_3 Bi $4f_{7/2} = 161.0$ eV and BiI_3 Bi $4f_{7/2} = 159.5$ eV.^[19] Therefore, for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ only bismuth in an oxidation state of 3+ is present. For $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]_{0.67}\text{Br}_{0.33}$, which is a solid at room temperature, four peaks are observed in the Bi 4f region, which corresponds to two electronic environments. The binding energies for the Bi $4f_{7/2}$ peaks are 159.7 eV and 157.1 eV, which are due to the ionic liquid and bismuth metal respectively (bismuth metal $4f_{7/2}$ reference = 157.1 eV).^[19] It is very likely that the bismuth metal was formed by X-ray damage (or low energy electrons produced by the X-rays). The observation that the binding energies are the same for the Bi $4f_{7/2}$ peak due to $[\text{BiBr}_x\text{A}_y]^-$ demonstrates that changing one ligand on the bismuth centre makes relatively little difference to the electronic environment of the bismuth atom.

Two peaks can be seen in Br 3d region for all four halobismuthate ionic liquids and also $[\text{C}_8\text{C}_1\text{Im}]\text{Br}$, $3d_{3/2}$ and $3d_{5/2}$ (Figure 4b). These two peaks are due to spin orbit coupling; therefore, for all five ionic liquids studied here there is one Br electronic environment only. The binding energy for the Br $3d_{5/2}$ peak is 67.5 eV for $[\text{C}_8\text{C}_1\text{Im}]\text{Br}$, the same as recorded within the error of the experiment (± 0.1 eV) by Villar-Garcia *et al.*^[18c] This peak is due to Br^- , and can be used as a reference. For all four halobismuthate ionic liquids, the binding energy of the Br $3d_{5/2}$ peak is 68.8 ± 0.1 eV, *i.e.* the same, within the error of the experiment (± 0.2 eV). The difference in binding energy between the Br $3d_{5/2}$

peak for Br^- and the $\text{Br } 3d_{5/2}$ peak for the four halobismuthate ionic liquids is ~ 1.3 eV. This difference is far larger than the error of the experiment (± 0.2 eV). Therefore, it can be conclusively stated that there is no lone Br^- present in any of the four bismuth-containing ionic liquids, demonstrating that all Br atoms in the halobismuthate ionic liquids are all ligands attached to Bi. For $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]_{0.67}\text{Br}_{0.33}$, the additional Br atoms are accommodated as ligands by the Bi. This observation is supported by crystallography, where anionic bismuth complexes are observed with large anions, such as $[\text{N}_{1,1,1,1}]_3[\text{Bi}_2\text{Cl}_9]$ and $[\text{N}_{1,1,1,1}]_3[\text{Bi}_2\text{Br}_9]$ (where $[\text{N}_{1,1,1,1}]^+$ = tetramethylammonium).^[20] In addition, the Br in the four halobismuthate ionic liquids was significantly more positively charged than Br^- . The Br^- anion has a charge of a little less than -1 e due to charge donation to the cation; from calculations on ion pairs, -0.8 e has been determined.^[21] This is to be expected given the Lewis acidity of Bi^{3+} . The charge of the Br atoms in the halobismuthate ionic liquids will be closer to 0 e than for Br^- . The observation that the $\text{Br } 3d_{5/2}$ peak due to $[\text{BiBr}_x\text{A}_y]^-$ is the same demonstrates that changing one ligand on the bismuth centre makes relatively little difference to the charge on the Br^- ligands.

For the Cl 2p region for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, two Cl peaks are observed, due to spin orbit coupling (Figure S15). Therefore, there is one Cl electronic environment for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$. The Cl $2p_{3/2}$ binding energy for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ is 198.5 eV. The Cl $2p_{3/2}$ binding energy for $[\text{C}_8\text{C}_1\text{Im}]\text{Cl}$ is 197.1 eV.^[18c] Therefore, there is no free Cl^- present in $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$. In addition, Cl in $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ is more positively charged than in $[\text{C}_8\text{C}_1\text{Im}]\text{Cl}$. For the I 3d region for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$, two I peaks are observed, due to spin orbit coupling (Figure S17). Therefore, there is one I electronic environment for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$. The I $3d_{5/2}$ binding energy for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ is 619.5 eV. The I $3d_{5/2}$ binding energy for $[\text{C}_8\text{C}_1\text{Im}]\text{I}$ is 618.5 eV.^[18c] Therefore, there is no free I^- present for $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$. In addition, I in $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ is more positively charged than in $[\text{C}_8\text{C}_1\text{Im}]\text{I}$.

For all five ionic liquids studied here, only one nitrogen peak is observed (Figures S15-S19), as expected for $[\text{C}_n\text{C}_1\text{Im}]^+$, as the two nitrogen atoms are indistinguishable by XPS.^[17] For the three halobismuthate compounds that are liquid at room temperature, $\text{N}_{\text{cation}} 1s$ is 401.9 eV. Due to the relatively large error for the halobismuthate ionic liquids (± 0.2 eV), the binding energies, whilst all intermediate between $\text{N}_{\text{cation}} 1s$ for $[\text{C}_8\text{C}_1\text{Im}][\text{Tf}_2\text{N}]$ and $[\text{C}_8\text{C}_1\text{Im}]\text{Br}$, cannot be distinguished.

2.7 Speciation

Due to the labile nature of the bismuth-halogen bond, it is expected that a number of different anionic species might form in bismuth halometallates, as with haloaluminates.^[7]

Mass spectra (section 2.1) indicate that when more than one halide element is present, the ligands are randomly distributed amongst the bismuth ions. That is to say, when chloride and bromide are both present, all possible ligand combinations are found (e.g. $[\text{BiBr}_4]^-$, $[\text{BiClBr}_3]^-$, $[\text{BiCl}_2\text{Br}_2]^-$ etc.). The existence of mixed halide dialkylimidazolium haloaluminates, specifically $[\text{AlBrCl}_3]^-$ and $[\text{AlCl}_3\text{I}]^-$, has previously been reported by Koronais and Osteryoung.^[22] Based on the MALDI⁻ m/z peak intensities, the most common complex in both $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ is not the stoichiometric ratio of ions (i.e. $[\text{BiBr}_3\text{Cl}]^-$ or $[\text{BiBr}_3\text{I}]^-$) but $[\text{BiBr}_4]^-$. This suggests that there might be a preference for the ions to disproportionate, forming $[\text{BiBr}_4]^-$ and $[\text{BiBr}_2\text{Cl}_2]^-$ rather than two $[\text{BiBr}_3\text{Cl}]^-$ ions.

The presence of uncoordinated halide ions is unclear based on the available results. Peaks for uncoordinated halide ions (Br^- and I^-) were present in the LSIMS⁻ mass spectra, although these may

predominantly have been the result of fragmentation of complex ions upon ionisation. XP measurements only detected the presence of one environment for Cl, Br and I, and the binding energies were different to those of free ions measured in bismuth-free reference compounds. Even when excess Br⁻ was present in the compound [C₄C₁Im][BiBr₄]_{0.67}Br_{0.33}, only one Br environment was observed suggesting that the extra ions were bound to the bismuth ions. On the other hand, because T_{onset} for [C₄C₁Im][BiCl₄] was close to T_{onset} [C₄C₁Im]Cl (see section 2.3), the chloride ligands in the [BiCl₄]⁻ anion may be more labile and free to attack the dialkylimidazolium cation than in the other halobismuthate analogues. This is in contrast to [C₄C₁Im][BiBr₄] which has a T_{onset} of 380 °C; much higher than the T_{onset} of 273 °C for [C₄C₁Im]Br.^[23] Furthermore, since iodide is more nucleophilic than chloride and bromide, and based on the results of Chan *et al.* demonstrating the thermolysis of [C₄C₁Im]I at 220 °C to 260 °C, the T_{onset} of [C₄C₁Im]I should be expected to be similar (no higher) than that of [C₄C₁Im]Cl or [C₄C₁Im]Br.^[24] Given that the T_{onset} of [C₄C₁Im][BiBr₃I] is 372 °C, the iodide ion appears to be strongly bound to the bismuth centre.

When the ratio of bismuth halide to dialkylimidazolium is 2:1 (*i.e.* [C₄C₁Im][Bi₂Br₇] and [C₄C₁Im][Bi₂BrI₆]), LSIMS⁻ peaks for bismuth dimers (*e.g.* [Bi₂Br₃]⁻ and [Bi₂I₃]⁻) and even trimers ([Bi₃Br₆I₃]⁻) are seen. The relative abundance of these peaks is low (2% to 5%) while monomeric species still dominate. It is unclear whether this is because the dimers readily fragment upon ionisation or whether their relative abundance in the liquid is similarly low. XP spectra only detect the presence of one bismuth environment and are not able to distinguish the presence of differently coordinated bismuth centres or bismuth dimers in a single liquid.

3 Experimental

3.1 Synthesis of halobismuthate compounds

All ionic liquids were prepared in approximately 10 g quantities by mixing a dialkylimidazolium halide salt with a bismuth halide salt in the desired molar ratio. The halide ion from the imidazolium salt transferred onto the bismuth salt to form the halobismuthate ion. The reagents were stirred under vacuum (<10⁻¹ mbar) at 100 °C until a homogeneous liquid had formed (5–30 min). For example: Dry [C₄C₁Im]Cl (2.92 g, 167 mmol) was added to BiBr₃ (7.50 g, 167 mmol) and stirred under vacuum at 100 °C for 30 min yielding a homogenous and transparent yellow liquid (10.42 g). Mixtures with [C₄C₁Im][NTf₂] were prepared by mixing [C₄C₁Im][NTf₂] with a pre-prepared imidazolium halobismuthate liquid in a 50:50 mol% ratio. [C₄C₁Im][Cl] (≥98.0%, Aldrich), [C₄C₁Im][Br] (>97.0%, Fluka), [C₂C₁Im][Cl] (98%, Aldrich), [C₄C₁Im][NTf₂] (≥98.0%, BASF), BiCl₃ (≥98%, Aldrich), BiBr₃ (≥98%, Aldrich) and BiI₃ (99.999%, Alfa Aesar) were dried under vacuum at 50 °C for several hours before use. The samples were stored in a vacuum desiccator prior to further study.

3.2 Density

Density measurements were made gravimetrically using a 1 ml volumetric flask calibrated against ultrapure water.

3.3 Viscosity

Steady shear rheology experiments were performed on a stress-controlled shear rheometer (T.A. Instruments, model AR-G2), with a 40 mm diameter cone-plate with angle 2° 0' 11" and a gap of 52 µm. The viscosity was measured over a range of temperatures, controlled by the inbuilt Peltier plate.

3.4 Mass spectrometry

ESI mass spectrometry was performed on a Waters 'LCT Premier' mass spectrometer. LSIMS mass spectrometry was performed on a 'Micromass AutoSpec Premier' mass spectrometer. MALDI mass spectrometry was performed on a Micromass 'MALDI-ToF' mass spectrometer.

3.5 NMR spectroscopy

NMR spectra were recorded on Bruker Avance-400 (¹H (400 MHz), ¹³C (100 MHz)) NMR spectrometers. Chemical shifts are reported in ppm (relative to the DMSO-d6 residual peak).

3.6 Thermo gravimetric analysis (TGA)

TGA spectra were obtained on a PerkinElmer 'Pyris 1 TGA' Thermogravimetric Analyzer, using platinum sample pans of 6 mm diameter. Temperature-ramped TGA experiments were carried out for ionic liquids in the range of 80 °C to 700 °C. Between 5 mg and 15 mg of the ionic liquid was measured into the platinum pan. A drying procedure was implemented: the ionic liquid was heated to 80 °C for 30 minutes in the TGA apparatus, in order to remove water. The drying procedure was justified since the onset decomposition temperature of each compound is substantially higher (at least 180 °C) than the 80 °C drying temperature, and the drying period is short. Therefore, actual decomposition of each ionic liquid during the drying period will be negligible. A ramping rate of 10 °C min⁻¹ and a nitrogen flow of 20 ml min⁻¹ were used for all temperature-ramped experiments.

3.7 Water content

Water content was measured using a Mettler Toledo V20 Volumetric Karl Fischer Titrator.

3.8 XPS

XPS was carried out using a Thermo K-alpha spectrometer utilising Al K α radiation (1486.6 eV) and a quartz crystal monochromator set in a 250 mm Rowland circle. The X-ray spot was focussed at the sample to a size of 400 µm. The base pressure was 10⁻⁹ mbar, and the analyser was a double focusing 180° hemisphere with mean radius 125 mm which was run in constant analyser energy mode. The pass energy was set to 200 eV for survey scans and 20 eV for high resolution regions. The detector was a 128 channel position sensitive detector. The energy scale of the instrument was regularly calibrated using a three point (Cu, Ag, Au) scale. Drops of the samples were placed directly onto a stainless steel plate. These were placed in a loadlock and the pressure reduced to 10⁻⁷ mbar by pumping down overnight if necessary. There was no significant outgassing or boiling of the samples during this process. After attaining the required pressure, samples were transferred to the analysis chamber. Where necessary, etching was carried out using a 500 eV Ar ion gun. Charge compensation was achieved using a dual beam flood gun which applies both electrons and low energy Ar⁺ ions to the sample.

4 Conclusions

Ionic liquids containing bismuth-centred halometallate anions have been synthesised as confirmed by NMR Spectroscopy, Electrospray Ionisation Mass Spectrometry, Liquid Secondary Ion Mass Spectrometry, Matrix-assisted Laser Desorption/Ionization Mass Spectrometry and XPS. Mass Spectrometry data indicated the existence of both monomeric and dimeric bismuth complexes, while XP spectra were able to identify one environment for Bi, Cl, Br and I. When more than one halide species was present in the liquid, the halide ions appeared to be distributed randomly between bismuth ion centres, resulting in the formation of many different anionic complexes.

The dialkylimidazolium halobismuthates have the highest densities yet reported for ionic liquids, due to the high molecular mass of bismuth and the complexed Br^- and I^- ions. The liquids have high viscosities but can be mixed with low viscosity $[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$, bringing about a dramatic reduction in viscosity. Furthermore, viscosity decreased rapidly upon heating. The majority of halobismuthate ionic liquids were shown to exhibit excellent thermal stability with T_{onset} decomposition temperatures upwards of 370 °C.

Given that bismuth is used extensively in synthetic chemistry, as are haloaluminates, it is likely that halobismuthate ionic liquids could be used in synthetic applications.^[6, 25] The room temperature electrochemical deposition of bismuth may also be possible. Various engineering applications are envisaged for these liquids in light of their combination of very high density, chemical stability, stability at elevated temperatures, low volatility, and moderate viscosity.

5 Acknowledgements

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Figures



Figure 1. $[\text{C}_4\text{C}_1\text{Im}]^+$ based halobismuthate ionic liquids with different anions (from left to right): $[\text{BiI}_3\text{Br}]^-$, $[\text{BiI}_3\text{Cl}]^-$, $[\text{BiBr}_2\text{I}_2]^-$, $[\text{BiCl}_2\text{I}_2]^-$, $[\text{BiBr}_2\text{ClI}]^-$, $[\text{BiBr}_4]^-$, $[\text{BiBr}_3\text{Cl}]^-$, $[\text{BiBr}_3\text{Cl}]^-_{0.5}[\text{NTf}_2]^-_{0.5}$, $[\text{BiBr}_2\text{Cl}_2]^-$, $[\text{BiBrCl}_3]^-$, $[\text{BiCl}_4]^-$.

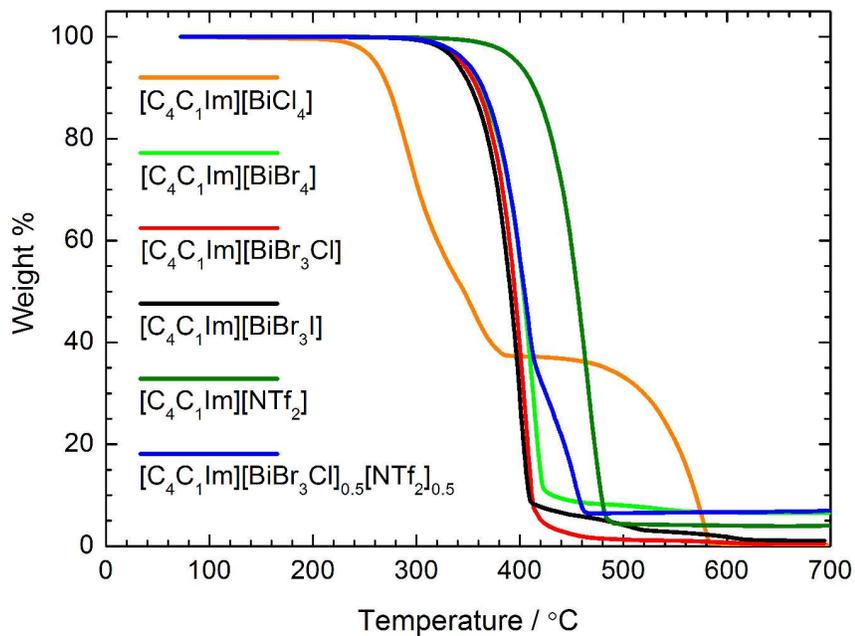
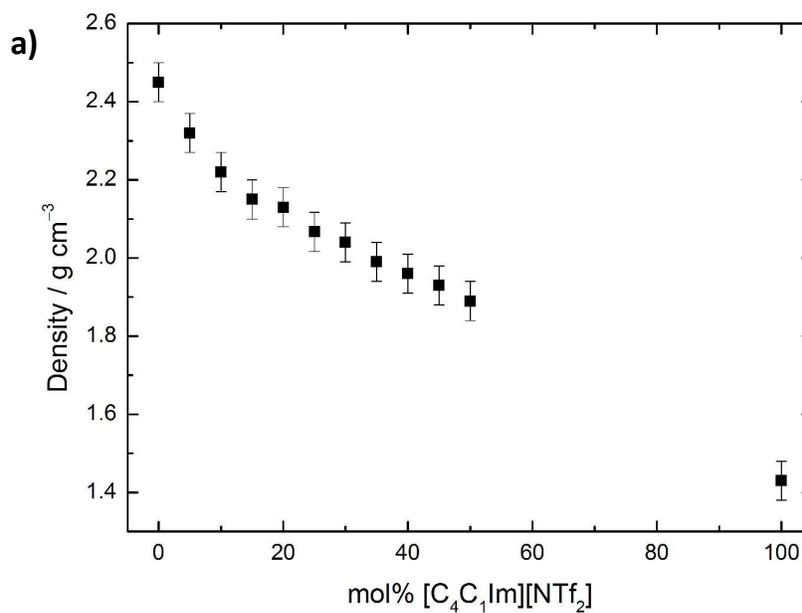


Figure 2. Temperature-ramped TGA thermographs of halobismuthate liquids, $[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$, and the mixture $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]_{0.5}[\text{NTf}_2]_{0.5}$, measured with a heating rate of $10\text{ }^\circ\text{C min}^{-1}$. As the temperature of the liquids is increased, the mass loss from evaporation of decomposition products is recorded simultaneously.



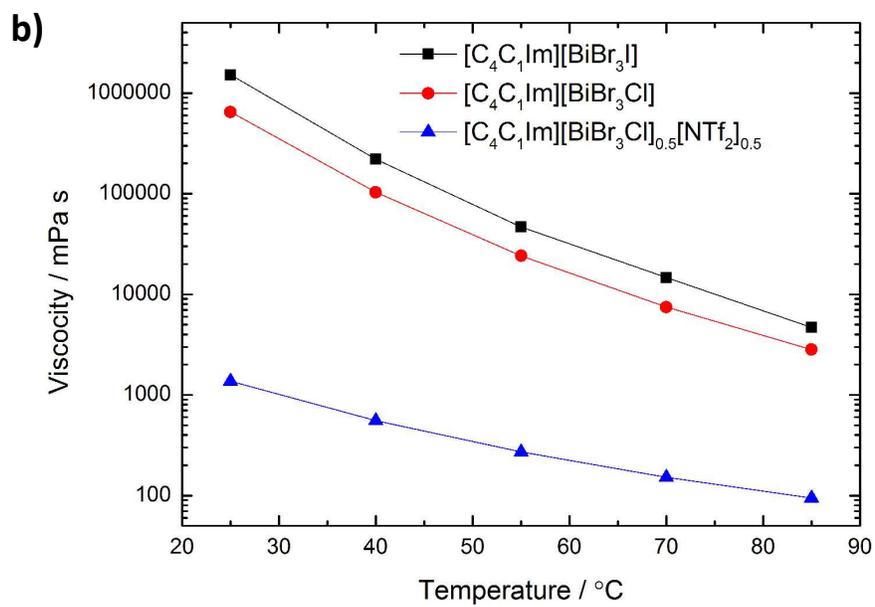


Figure 3. a) Density of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]/[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$ mixtures with varying $[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$ concentration. b) Viscosity of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]_{0.5}[\text{NTf}_2]_{0.5}$ between 25 °C and 85 °C. Error bars of $\pm 5\%$ are not visible.

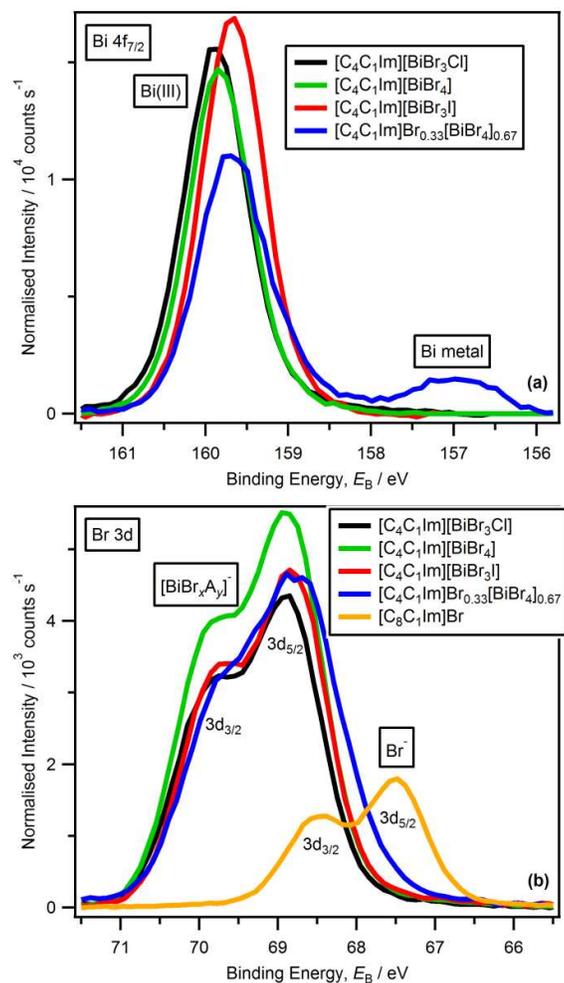


Figure 4. XP spectra for the four bismuth-containing ionic liquids ([C₄C₁Im][BiBr₃Cl], [C₄C₁Im][BiBr₄], [C₄C₁Im][BiBr₃I], and [C₄C₁Im][BiBr₄]_{0.67}Br_{0.33}) and [C₈C₁Im]Br: (a) Bi 4f_{7/2}, (b) Br 3d. The experimental values are derived from XP spectra taken at $\theta = 0^\circ$. XP spectra were charge corrected by referencing C_{alkyl} 1s to 285.0 eV. To aid visual interpretation of the XP spectra presented here, all XP spectra are normalised to the fitted area of N_{cation} 1s, the peak due to the nitrogen atoms in the imidazolium ring.

Tables

Ionic Liquid	Density $\pm 0.05 / \text{g cm}^{-3}$	Viscosity $\pm 5\% / \text{Pa s @ 25 } ^\circ\text{C}$	Glass transition/ $^\circ\text{C}$
$[\text{C}_2\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$	2.62*		<50
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$	2.41	648	<25
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$	2.49		<25
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}][\text{BiBr}_3]_{0.5}$	2.65		<25
$[\text{C}_4\text{C}_1\text{Im}][\text{Bi}_2\text{Br}_6\text{Cl}]$	2.92		<50
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]_{0.5}[\text{NTf}_2]_{0.5}$	1.89	1.370	<25
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3]$	2.63	1517	<25
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBrI}_3]$	2.78		<50
$[\text{C}_4\text{C}_1\text{Im}][\text{Bi}_2\text{BrI}_6]$	3.43**		<100

*measured at 50 $^\circ\text{C}$

**measured at 100 $^\circ\text{C}$

Table 1. Densities, viscosities and upper bound of glass transition temperature of a range of halobismuthate ionic liquids.

Ionic Liquid	$T_{\text{onset}} / ^\circ\text{C}$
$[\text{C}_4\text{C}_1\text{Im}][\text{NTf}_2]$	425
$[\text{C}_4\text{C}_1\text{Im}][\text{BiCl}_4]$	260
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$	375
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$	380
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3]$	372
$[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]_{0.5}[\text{NTf}_2]_{0.5}$	370

Table 2. Thermal decomposition temperatures, T_{onset} , for a range of halobismuthate ionic liquids.

		Cation				Anion				Contamination
		C _{N-C-N} 1s	C _{hetero} 1s	C _{alkyl} 1s	N _{cation} 1s	Cl _{anion} 2p _{3/2}	Br _{anion} 3d _{5/2}	I _{anion} 3d _{5/2}	Bi _{anion} 4f _{7/2}	Bi _{metal} 4f _{7/2}
	RSF	0.205	0.205	0.205	0.350	0.367	0.367	5.610	4.250	4.250
[C ₄ C ₁ Im][BiBr ₃ Cl]	Nominal	1	4	3	2	1	3	0	1	0
[C ₄ C ₁ Im][BiBr ₃ Cl]	Experimental	1.7	3.5	3.4	2.1	1.0	2.6		0.8	
[C ₄ C ₁ Im][BiBr ₄]	Nominal	1	4	3	2	0	4	0	1	0
[C ₄ C ₁ Im][BiBr ₄]	Experimental	1.4	3.9	3.3	2.2		3.5		0.8	
[C ₄ C ₁ Im][BiBr ₃ I]	Nominal	1	4	3	2	0	3	1	1	0
[C ₄ C ₁ Im][BiBr ₃ I]	Experimental	1.5	3.7	3.3	2.1		2.8	0.9	0.8	
[C ₄ C ₁ Im][BiBr ₄] _{0.67} Br _{0.33}	Nominal	1	4	3	2	0	3	0	0.67	0
[C ₄ C ₁ Im][BiBr ₄] _{0.67} Br _{0.33}	Experimental	1.3	3.4	3.8	1.8		2.7		0.5	0.1
[C ₈ C ₁ Im]Br	Nominal	1	4	7	2	0	1	0	0	0
[C ₈ C ₁ Im]Br	Experimental	1.3	4.0	6.9	1.9		0.9			

Table 3. Quantitative analysis of the XP spectra of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]_{0.67}\text{Br}_{0.33}$ and $[\text{C}_8\text{C}_1\text{Im}]\text{Br}$. The nominal and the experimentally determined composition in number of atoms are given for the various elements constituting the ionic liquids. The experimental values were derived from XP spectra taken at $\theta = 0^\circ$. The RSFs were taken from ref ^[14] and ^[15]. $\text{C}_{\text{N-C-N}}$ refers to the carbon atom at the C^2 position of the imidazolium ring; C_{hetero} refers to all other carbon atoms bonded to a heteroatom, in this case nitrogen (*i.e.* C^4 , C^5 , C^6 and C^7).

Sample	Cation				Anion				Contamination
	C ² 1s	C _{hetero} 1s	C _{alkyl} 1s	N _{cation} 1s	Cl _{anion} 2p _{3/2}	Br _{anion} 3d _{5/2}	I _{anion} 3d _{5/2}	Bi _{Bi(III)} 4f _{7/2}	Bi _{metal} 4f _{7/2}
[C ₄ C ₁ Im][BiBr ₃ Cl]	287.2	286.4	285.0	401.9	198.5	68.9		159.9	
[C ₄ C ₁ Im][BiBr ₄]	287.4	286.4	285.0	401.9		68.8		159.8	
[C ₄ C ₁ Im][BiBr ₃ I]	287.3	286.4	285.0	401.9		68.8	619.5	159.7	
[C ₄ C ₁ Im][BiBr ₄] _{0.67} Br _{0.33}	287.3	286.4	285.0	401.8		68.7		159.7	157.1
[C ₈ C ₁ Im]Br	287.2	286.3	285.0	401.6		67.4			
[C ₈ C ₁ Im]Cl (ref ^[18c])			285.0	401.7	197.1				
[C ₈ C ₁ Im]Br (ref ^[18c])			285.0	401.7		67.5			
[C ₈ C ₁ Im]I (ref ^[18c])			285.0	401.9			618.5		
Bi metal (ref ^[19])									157.1
BiF ₃ (ref ^[19])								161.0	
BiI ₃ (ref ^[19])								159.5	

Table 4. Binding energies in eV from the XP spectra of $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]_{0.67}\text{Br}_{0.33}$ and $[\text{C}_8\text{C}_1\text{Im}]\text{Br}$. The experimentally determined binding energy values are given for the various elements constituting the ionic liquids. The experimental values were derived from XP spectra taken at $\theta = 0^\circ$. XP spectra were charge corrected by referencing $\text{C}_{\text{alkyl}} 1s$ to 285.0 eV. For $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{Cl}]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]$, $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_3\text{I}]$ and $[\text{C}_4\text{C}_1\text{Im}][\text{BiBr}_4]_{0.67}\text{Br}_{0.33}$, the error was ± 0.2 eV; for $[\text{C}_8\text{C}_1\text{Im}]\text{Br}$, the error was ± 0.1 eV.

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