










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Study of the thermally-activated delayed fluorescence (TADF) mechanism of phenothiazine–dibenzothiophene-*S,S*-dioxide electron donor–acceptor dyads using steady-state and time-resolved optical and electron paramagnetic resonance spectroscopies

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We studied a series of thermally activated delayed fluorescence (TADF) emitters based on phenothiazine–dibenzothiophene-*S,S*-dioxide (PTZ-DTO) electron donor–acceptor (D–A) dyads, using femtosecond/nanosecond transient absorption (fs/ns-TA) spectroscopy and pulsed laser-excited time-resolved electron paramagnetic resonance (TREPR) spectroscopy. In the analyzed compounds, a Se atom replaced the S atom in the phenothiazine (PTZ) unit to study the heavy atom effect on the reverse intersystem crossing (RISC). Moreover, oxidation of the PTZ to sulfoxide and sulfone was also used to tune the energy of the charge separated (CS) state, while keeping the other factors (³LE state energy, LE = locally excited, and electronic coupling between the donor and acceptor) intact to large extent. Fs-TA spectra show that charge separation occurs rapidly (ca. 7.6 ps) in non oxidized compounds, while for the compounds with the PTZ unit oxidized the CS is slightly slower (10.9 ps). Ns-TA spectroscopy demonstrated the coexistence of ³CS and ³LE states for the non oxidized dyads and the absence of heavy atom effect on RISC. The oxidation of the PTZ unit increased the CS state energy, so that only the ³LE state remains observed. All our results demonstrated that RISC rate constants are not enhanced in the presence of heavy atoms for the studied dyads. TREPR spectra show the presence of the ³LE state, and that the triplet state is formed most likely *via* spin–orbit charge transfer intersystem crossing (SOCT-ISC) because the electron spin polarization (ESP) phase pattern is (e, a, e, a, e, a). These studies are useful for an in-depth understanding of the photophysics of the electron D–A TADF emitters.

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Introduction

Electron donor–acceptor (D–A) dyads showing thermally-activated delayed fluorescence (TADF) properties have attracted notable attention because of their potential application in organic light emitting diodes (OLED).^{1–16} The most striking advantage of these compounds is that although mostly not containing heavy atoms (*e.g.*, Ir, Pt, Br, I, *etc.*), they are able to harvest triplet excitons and to populate the emissive singlet excited state *via* reverse intersystem crossing (RISC), thus emitting both prompt as well as the delayed fluorescence.^{1,2} The development of new TADF emitters and their application in OLED devices are very hot topics in current research, although a detailed understanding of the mechanisms of TADF, ISC, and RISC is missing in most cases.^{4,5}

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Reverse ISC and delayed fluorescence were initially explained in terms of a two-state model,^{17–20} according to which the triplet excited state (T_1) and singlet excited state (S_1) should have similar energy, such as to minimize the electron exchange energy (J) and the S_1/T_1 energy gap ($\Delta E_{S_1/T_1}$), thus enhancing RISC, finally leading to TADF. A small $\Delta E_{S_1/T_1}$ is beneficial for both ISC and RISC,²¹ but the electronic configuration of the involved S_1 and T_1 states are not unambiguously elucidated. In most cases the two states are assumed to be charge separated (CS) states and since the ${}^3\text{CS} \rightarrow {}^1\text{CS}$ transition is quantum mechanically forbidden,^{22,23} a simple two states model is often non-sufficient to rationalize the observation of delayed fluorescence. Recently, we have confirmed this postulation with several experiments.^{24–27} It was also proposed that the activation energy for the RISC is more important than the energy gap between the S_1 and T_1 states for observing TADF.²⁸

In most cases, TADF emitters were only studied using steady state and time-resolved photoluminescence spectroscopies, but these techniques don't give direct experimental evidence for the existence of dark states such as ${}^3\text{CS}$ and ${}^3\text{LE}$ states (LE: locally excited).

Recently, it has been proposed that TADF should be rationalized in the frame of a three-state model,^{23,29–37} involving an emissive ${}^1\text{CS}$ state, a dark ${}^3\text{CS}$ state and a chromophore-localized ${}^3\text{LE}$ dark state. The latter ${}^3\text{LE}$ state is necessary to mediate the transition from the ${}^3\text{CS}$ state to the emissive ${}^1\text{CS}$ state (eqn (1)).^{29,30}

$$k_{\text{RISC}} = \frac{2\pi}{\hbar} \left| \frac{\langle \psi_{1\text{CS}} | \hat{H}_{\text{SOC}} | \psi_{3\text{LE}} \rangle \langle \psi_{3\text{LE}} | \hat{H}_{\text{vib}} | \psi_{3\text{CS}} \rangle}{E_{3\text{CS}} - E_{3\text{LE}}} \right|^2 \delta(E_{1\text{CS}} - E_{3\text{LE}}) \quad (1)$$

Theoretical studies have shown that for a well-known TADF emitter, **PTZ-DBTO2**, ISC between the ${}^1\text{CS}$ and the ${}^3\text{CS}$ state is driven by a hyperfine interaction of 0.2 cm^{-1} , whereas the spin-orbit coupling between the singlet charge separated (${}^1\text{CT}$) state and the ${}^3\text{LE}$ state is 2 cm^{-1} , and the vibronic coupling between the ${}^3\text{CT}$ and the ${}^3\text{LE}$ is 65 cm^{-1} .²⁹ A theoretical study based on the spin-vibronic model predicts that the efficient RISC of **PTZ-DBTO2** is due to the stronger S_1/T_2 coupling as compared to that between the S_1/T_1 states.³² However, experimental evidence confirming the co-existence of these three states is still scarce. Since the nature of the CS state of an electron D–A dyad depends on the electron dipolar interaction magnitude, the ion pairs can be considered either as free radical pairs (if the separation between the anion and the cation is large, e.g., $r > 30 \text{ \AA}$) or as spin-correlated radical pairs (SCRPs, where $r \sim 10\text{--}20 \text{ \AA}$), or finally as ${}^3\text{CS}$ state (with short separation between the anion and cation, for instance by a single bond, in this case the J value is large). All these species present similar bands in transient optical absorption spectra, which makes it difficult, if not impossible, to unambiguously discriminate between them only relying on optical spectroscopic methods.²⁵

Recently femtosecond transient absorption (fs-TA) spectroscopy and nanosecond transient absorption (ns-TA) spectroscopy were used to characterize the photophysical processes of TADF

emitters, evidencing the formation of ${}^1\text{LE}$ or ${}^1\text{CS}$ states as well as the ISC process.^{25,38–42} Experiments have shown that the rate of RISC can be significantly enhanced if the energy of ${}^3\text{LE}$ state is similar to that of the ${}^3\text{CS}$ state, thus contributing to efficient TADF.^{24,43} In contrast, the RISC process slows down significantly when ${}^3\text{LE}$ and ${}^1\text{CS}$ states with similar energies are missing, making TADF less efficient.^{26,44,45} Both fs-TA and ns-TA spectroscopies have provided direct experimental evidence for the spin-vibronic coupling mechanism,^{29,30} being able to detect dark states such as ${}^3\text{CS}$ and ${}^3\text{LE}$ states.

A discrimination between ${}^1\text{CS}$ and ${}^3\text{CS}$ states can be made based on their lifetime and emissive properties. In previous studies, we have confirmed that if the energy gap between the ${}^3\text{CS}$ and ${}^3\text{LE}$ states is large, the ${}^3\text{CS} \rightarrow {}^1\text{CS}$ transition is too slow to induce any TADF.^{24,26,27} Another useful technique, able to selectively detect the transient paramagnetic species involved in the TADF processes, such as the ${}^3\text{LE}$ and ${}^3\text{CS}$ states, as well as their interconversion and the ISC mechanism, is the pulsed laser excited time-resolved electron paramagnetic resonance (TREPR) spectroscopy.^{46,47} For instance, using TREPR spectroscopy it was proposed that for the carbazole-derived TADF compounds (**4Cz-IPN**), ISC occurs through a radical pair ISC (RP-ISC) mechanism, based on the electron spin polarization (ESP) of the ${}^3\text{LE}$ state.⁴⁸ Another report, analyzing the 4-(3,6-di-*tert*-butyl-9*H*-carbazol-9-yl)diphenyl sulfone (**DTCz-DPS**) TADF emitter through TREPR spectroscopy, showed that vibrations play a critical role for ISC.⁴⁹ Evans *et al.* also found that a prominent ${}^3\text{CS}$ state is present in case of good TADF emitters. Generally, the ${}^3\text{CS}$ state has a significantly smaller zero field splitting (ZFS) D parameter than the ${}^3\text{LE}$ state, because the average distance between the electrons is usually larger (point-dipole approximation).^{43,48} We used TREPR spectroscopy to study the photophysical mechanism of the TADF process, observing for instance both a ${}^3\text{LE}$ state and a ${}^3\text{CS}$ state in a naphthalimide–phenothiazine (**NI-PTZ**) dyad, which could be distinguished on the basis of their electron spin polarization pattern.⁴³ We observed that the electron spin dipolar interaction is much stronger for ${}^3\text{CS}$ states compared to SCRPs states.^{43,50,51} This arises from the short distance between the electron donor and acceptor in **NI-PTZ** dyads,^{49,52} allowing us to observe the interconversion between the ${}^3\text{CS}$ and ${}^3\text{LE}$ states, which is



Scheme 1 Molecular structures of the TADF emitters and reference compounds used in the study. The red units are the electron donors, and the blue units are the electron acceptors.



a strong direct experimental evidence for the spin-vibronic coupling model of TADF.

Although some evidence is present in the literatures, more experimental studies are needed to verify the above observations. Herein we selected a series representative TADF emitters based on the electron D–A couple phenothiazine–dibenzothiophene-*S,S*-dioxide (**PTZ-DTO**, Scheme 1).^{53,54} The photophysical properties of the compounds were studied by using steady state UV/vis absorption and fluorescence spectroscopy, and their excited state dynamics was analysed with fs-TA and ns-TA spectroscopy. Furthermore, TREPR spectroscopy was used for characterizing the transient paramagnetic species and their electron spin dynamics.

Results and discussion

Molecular structure designing rationales

We selected the D–A TADF emitters **PTZ-DTO** and **DPTZ-DTO**, based on phenothiazine (PTZ) as the electron donor, because of its strong electron-donating ability ($E_{\text{OX}} = 0.18 \text{ V}$, vs. Fc/Fc^+),⁵⁵ and dibenzothiophene-*S,S*-dioxide (DTO) as the electron acceptor. The TADF properties of these two compounds have been previously reported and their application in polymer light emitting diodes OLED devices has been studied.^{53,56} However, details on the ISC and RISC mechanisms were not reported.

We then tuned the redox potential of the compounds by oxidizing the sulfur atom into sulfoxide and sulfone (**PTZ-O-DTO**, **PTZ-O2-DTO**, **DPTZ-O-DTO**), changing the electron-donating ability of PTZ, thus altering the energy of the CS state and consequently the TADF properties.⁵⁷ We also substituted the sulfur (S) atom in the PTZ unit with selenium (Se), to change the photophysical properties of the dyads due to the introduction of a heavy atom (**PSeZ-DTO**).⁵⁸ Previous studies have shown that the introduction of Se atoms in TADF emitters enhance the RISC, which is beneficial for OLED devices.^{58–61} The molecular structure of the compounds has been confirmed by ^1H NMR, ^{13}C NMR and HRMS characterizations.

UV/vis absorption and fluorescence spectra

The UV/vis absorption spectra of the compounds in DCM are presented in Fig. 1. All compounds absorb in the 300–420 nm



Fig. 1 UV/vis absorption spectra of the compounds in dichloromethane (DCM). $c = 3.0 \times 10^{-4} \text{ M}$, $25 \text{ }^\circ\text{C}$.

spectral range. The absorption spectra of PTZ and DTO have been previously reported, and their absorption bands are centered at *ca.* 320 nm.⁵⁴ Since the HOMO and LUMO molecular orbitals have poor spatial overlap (SI), the absorption bands of **PTZ-DTO** and **DPTZ-DTO** result from the superposition the PTZ and DTO absorption bands, indicating negligible interaction between the electron donor and acceptor at the ground state. The absorption spectrum of **PSeZ-DTO** is similar to that of **PTZ-DTO**, although the absorption band centered at 320 nm is slightly stronger. For **PTZ-DTO**, **PSeZ-DTO** and **DPTZ-DTO**, a weak CT shoulder absorption band was observed. For **PTZ-O-DTO**, **PTZ-O2-DTO** and **DPTZ-O-DTO**, the absorption peaks slightly shift but remain in the 300–375 nm spectral range. For **PTZ-O-DTO** and **PTZ-O2-DTO** the CT absorption is hardly observed, even at high concentrations. It has previously been reported that due to the weak yet finite coupling between the DTO and PTZ moieties, there is an absorption band centered at 450 nm, assigned to the $\text{S}_0 \rightarrow {}^1\text{CT}$ transition.^{53,62,63} For **DPTZ-O-DTO**, the CT absorption band is blue shifted and becomes weaker, indicating that the coupling between the electron donor and acceptor is weaker after oxidation of the PTZ unit. We confirmed that these compounds are photostable with continuous photo-irradiation experiments.⁶⁴

The fluorescence spectra of the compounds were studied (Fig. 2). The spectra of **PTZ-DTO** and **DPTZ-DTO** are similar to those reported previously.^{53,62,63} For all the three compounds **PTZ-DTO**, **PSeZ-DTO** and **DPTZ-DTO**, a broad and structureless emission band in the range of 500–800 nm is observed, attributed to the ${}^1\text{CT} \rightarrow \text{S}_0$ emission. However, the fluorescence emission wavelength and intensity of the three compounds are different. For **PTZ-DTO**, the fluorescence band is centered at 568 nm (Fig. 2a), for **PSeZ-DTO** it is centered at 559 nm (Fig. 2b), while it is centered at 595 nm for **DPTZ-DTO** (Fig. 2c).

We then measured the fluorescence of the compounds under different atmospheres, finding that it is significantly enhanced under N_2 atmosphere and quenched by O_2 under air atmosphere, indicating the involvement of triplet states in the emission process.^{65,66} Greatly enhanced fluorescence under N_2 is typical for compounds showing TADF.¹ The luminescence of **PTZ-O-DTO**, **PTZ-O2-DTO** and **DPTZ-O-DTO**, in which the PTZ unit is oxidized, were not significantly quenched under air atmosphere (Fig. S8), suggesting that the TADF properties of the oxidized compounds is not significant.^{26,67}

We also measured the fluorescence spectra of the compounds in different solvents (Fig. 2d–f). For **PTZ-DTO**, **PSeZ-DTO** and **DPTZ-DTO**, the emission bands are red-shifted with increased solvent polarity. The emission intensity in TOL is stronger than that in hexane (HEX), which is often observed in case of CS state emission.^{65,66,68,69} The fluorescence bands of **PTZ-O-DTO**, **PTZ-O2-DTO** and **DPTZ-O-DTO** in different solvents are also significantly red-shifted, but the fluorescence intensity is not significantly quenched, indicating that the emissive state might not be a conventional intramolecular charge transfer state (Fig. S8).





Fig. 2 Fluorescence emission spectra of the compounds under N₂ and air atmosphere of (a) PTZ-DTO, (b) PSeZ-DTO and (c) DPTZ-DTO in toluene (TOL). Fluorescence emission spectra of the compounds in different solvents of (d) PTZ-DTO, (e) PSeZ-DTO and (f) DPTZ-DTO in aerated solution. Optically-matched solutions were used, $A = 0.1$, $\lambda_{\text{ex}} = 310$ nm, 25 °C.

Fluorescence lifetime studies

The luminescence lifetime of the TADF emitters in TOL under air and N₂ atmospheres are compared in Fig. 3. The lifetime of the CS emission was monitored at 560 nm (Fig. 3a) for **PTZ-DTO** under N₂ atmosphere, and it displayed a biexponential decay, with two lifetimes of 22.8 ns (96.6%) and 1.4 μ s (3.4%), respectively. The short-lived component is assigned as prompt fluorescence, while the long-lived component as delayed

fluorescence.¹ In aerated solution, the lifetime of the delayed fluorescence is significantly shortened (45.5 ns), indicating that the triplet state is involved in the luminescence.^{2,70-72} The fluorescence decay of **PTZ-DTO** in the solid matrix Zeonex® COP 480R has been previously reported, also showing a double exponential decay.⁵³

Fig. 3b presents the fluorescence lifetime of **PSeZ-DTO** monitored at 550 nm. Similar to **PTZ-DTO**, the emission of **PSeZ-DTO** presents a biexponential decay in N₂ atmosphere: the

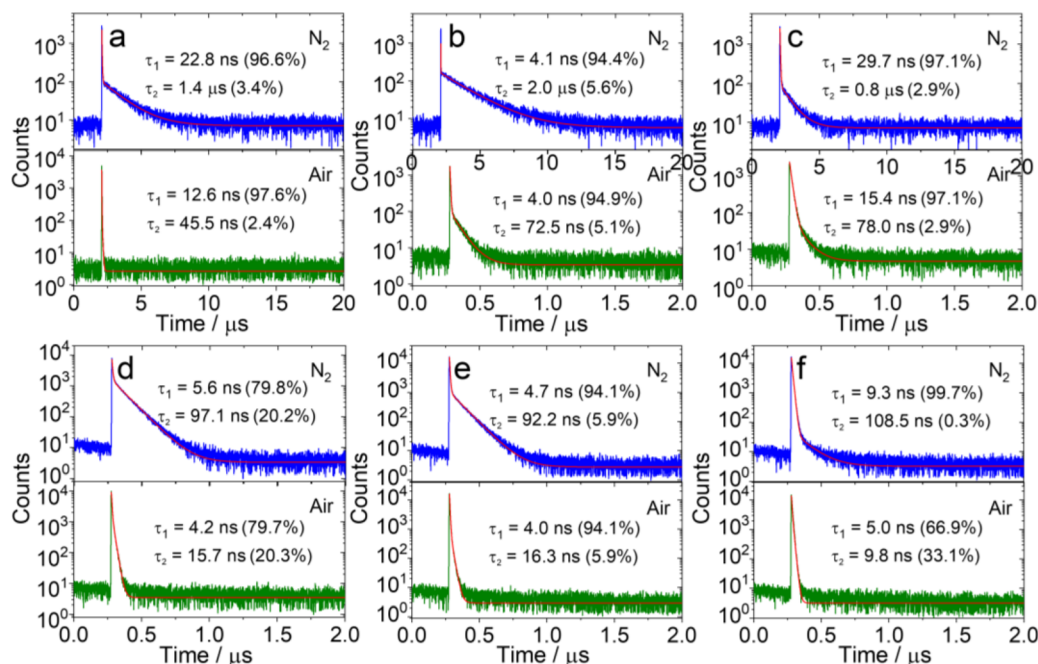


Fig. 3 Fluorescence decay traces of (a) PTZ-DTO at 560 nm, (b) PSeZ-DTO at 550 nm, (c) DPTZ-DTO at 580 nm, (d) PTZ-O-DTO at 400 nm, (e) PTZ-O2-DTO at 410 nm and (f) DPTZ-O-DTO at 470 nm under different atmospheres (N₂ and air). $\lambda_{\text{ex}} = 340$ nm, $c = 1.0 \times 10^{-5}$ M in TOL, 25 °C.



Table 1 Photophysical data of the compounds

Compounds	$\lambda_{\text{abs}}/\text{nm}^a$	ϵ^b	$\lambda_{\text{em}}/\text{nm}^c$	$\tau_{\text{F}}/\text{ns}^d$	Φ_{F}^e (%)	Φ_{Δ}^f (%)
PTZ-DTO	319	0.5	568	22.8 (96.6%) 1400 (3.4%)	6.1	0.2
PSeZ-DTO	322	0.6	559	4.1 (94.4%) 2000 (5.6%)	3.2	12.1
DPTZ-DTO	324	1.0	595	29.7 (97.1%) 800 (2.9%)	6.5	10.0
PTZ-O-DTO	325	0.6	439	5.6 (79.8%) 97.1 (20.2%)	4.5	59.0
PTZ-O2-DTO	321	0.7	405	4.7 (94.1%) 92.2 (5.9%)	5.7	69.2
DPTZ-O-DTO	330	1.1	470	9.3 (99.7%) 108.5 (0.3%)	5.6	64.6

^a Maximal UV/vis absorption wavelength, $c = 3.0 \times 10^{-4}$ M, 25 °C.

^b Molar absorption coefficient at absorption maxima, $\epsilon: 10^4 \text{ M}^{-1} \text{ cm}^{-1}$, in DCM. ^c Emission wavelength. ^d Fluorescence lifetime, $\lambda_{\text{ex}} = 340$ nm. ^e Fluorescence quantum yields, $\lambda_{\text{ex}} = 310$ nm.

^f Singlet oxygen quantum yields, Ru(bpy)₃[PF₆]₂ was used as standard compound ($\Phi_{\Delta} = 57\%$ in DCM), $\lambda_{\text{ex}} = 343$ nm.

lifetime of prompt fluorescence is 4.1 ns (94.4%) and that of delayed fluorescence is 2.0 μs (5.6%). Contrarily to what observed previously,^{58,61,73} in our case the presence of a heavy atom does not shorten the lifetime of the delayed fluorescence, suggesting that the heavy atom effects don't enhance significantly the rate of RISC in this compound. For **PSeZ-DTO**, the delayed fluorescence lifetime is shortened to 72.5 ns in aerated solution. We attribute the lack of heavy atom effect on the TADF property to the varied CS state energy upon the one-atom alternation of the molecular structure (see the Electrochemistry section).

The fluorescence decay traces of **DPTZ-DTO** is presented in Fig. 3c. Compared with **PTZ-DTO**, we observed a decreased lifetime for the slow decaying phase, which results 0.8 μs in the deaerated solution, while the lifetime of the fast decay component is 29.7 ns, longer than that of the **PTZ-DTO**. In this case the long lifetime is shortened under air atmosphere, indicating that the compound has TADF properties. The fluorescence of this compound behaves similarly to that of **DPTZ-DTO** ($\tau_{\text{PF}} = 33.0$ ns and $\tau_{\text{DF}} = 1.1 \mu\text{s}$).⁷⁴

Table 2 Singlet oxygen quantum yields (Φ_{Δ} in %) of the compounds in different solvents^a

Compounds	HEX	TOL	DCM	ACN	MeOH
PTZ-DTO	11.7	0.2	— ^b	— ^b	— ^b
PSeZ-DTO	9.6	12.1	— ^b	— ^b	— ^b
DPTZ-DTO	0.3	10	— ^b	— ^b	— ^b
PTZ-O-DTO	28.7	59.0	86.5	70.8	11.7
PTZ-O2-DTO	4.8	69.2	99.2	99.5	41.7
DPTZ-O-DTO	— ^b	64.6	35.2	53.8	0.6

^a The E_{T} (30) values of the solvents are 31.0 kcal mol⁻¹ (HEX), 33.9 kcal mol⁻¹ (TOL), 40.7 kcal mol⁻¹ (DCM), 45.6 kcal mol⁻¹ (ACN) and 55.4 kcal mol⁻¹ (methanol (MeOH)), respectively. Φ_{Δ} with Ru(bpy)₃[PF₆]₂ as standard ($\Phi_{\Delta} = 0.57$ in DCM) in different solvents, $\lambda_{\text{ex}} = 330$ nm.

^b Not observed.

We also studied the luminescence lifetimes of **PTZ-O-DTO** and **PTZ-O2-DTO** (Fig. 3d and e), finding also in this case a bi-exponential decay. Under N₂ atmosphere, the delayed fluorescence lifetime was determined 97.1 ns for **PTZ-O-DTO** and 92.2 ns for **PTZ-O2-DTO**, and both compounds displayed a high singlet oxygen quantum yield (Φ_{Δ} , Tables 1 and 2). For **DPTZ-O-DTO** (Fig. 3f), the delayed fluorescence lifetime component accounts for very little and is almost negligible.

The photophysical data of all compounds are summarized in Table 1. The fluorescence quantum yields (Φ_{F}) of the compounds are low, generally between 3 and 6%, indicating that non-radiative relaxation channels are quite efficient for all of them in fluid solution. We also measured the Φ_{Δ} of the compounds in different solvents to characterize the ISC efficiency. For **PTZ-DTO**, **PSeZ-DTO** and **DPTZ-DTO** we observed production of ¹O₂ only in low polar solvents, possibly because in polar media the CS state energy becomes lower than that of the ³LE state, thus preventing ¹O₂ production. Upon oxidation of the PTZ unit, the Φ_{Δ} of all compounds increases, especially in highly polar solvents. In this case the lowest energy triplet state is ³LE, according to the results obtained with ns-TA spectroscopy (see later section).

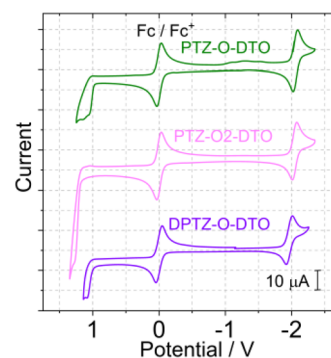


Fig. 4 Cyclic voltammograms of the compounds **PTZ-O-DTO**, **PTZ-O2-DTO** recorded in deaerated acetonitrile (ACN), **DPTZ-O-DTO** was studied in deaerated DCM. Ferrocene (Fc) was used as internal reference (set as 0 V in the cyclic voltammograms), 0.10 M Bu₄NPF₆ as supporting electrolyte. Scan rates: 50 mV s⁻¹, $c = 1.0 \times 10^{-4}$ M, 25 °C.

Table 3 Redox potentials of the compounds in different solvents^a

Compounds	E_{OX} (V)	E_{RED} (V)
PTZ-DTO ^b	+0.37	-2.12
PSeZ-DTO ^b	+0.45	-2.17
DPTZ-DTO ^c	+0.41	-2.15
PTZ-O-DTO ^b	+1.14	-2.05
PTZ-O2-DTO ^b	+1.28	-2.04
DPTZ-O-DTO ^c	+1.09	-1.96

^a Cyclic voltammetry in N₂-saturated solvents containing 0.10 M Bu₄NPF₆, glassy carbon electrode as the working electrode, Pt electrode as the counter electrode, Ag/AgNO₃ couple as the reference electrode and ferrocene (Fc/Fc⁺) as internal reference (set as 0 V in the cyclic voltammograms). ^b Measured in ACN. ^c Measured in DCM.



Electrochemical properties

The redox potentials of the compounds were measured with cyclic voltammetry (Fig. 4) and the data are compiled in Table 3. For the compounds **PTZ-O-DTO**, **PTZ-O2-DTO** and **DPTZ-O-DTO** the oxidation waves are irreversible and significantly stronger if compared with the unoxidized compounds, because of the weakened electron donating ability of the PTZ moiety after oxidation.^{55,57,75} The free energy changes for the intramolecular electron transfer process can be calculated using the Rehm-Weller equation. The energies of the CS states (E_{CS}) can be calculated with the eqn (2).⁷⁶⁻⁷⁸

A negative Gibbs free energy change (ΔG_{CS}) of the compounds in polar solvents indicates thermodynamically allowed charge separation.^{66,76,79-81} In the non-polar solvent *n*-hexane CS is thermodynamically inhibited. As the polarity of the solvent increases, the CS state energy decreases.⁸² As noticed from Table 4, the energy of the CS state increases after the oxidation of PTZ but still decreases with the increasing solvent polarity. Based on the electrochemical and fluorescence emission data, we found that the energy of the CS state changed after the introduction of Se atom, resulting in changes in the S_1 and T_1 energy gaps. The change in the energy level of the CS state reasonably explains why Se substitution failed to significantly improve RISC. The energy of the 3LE state of the compounds can be obtained from the 77 K phosphorescence spectrum. The 3LE state energy of **PTZ-DTO** and **PSeZ-DTO** is approximated as 2.71 eV. It was found the 3CS and 3LE states have similar energy in toluene (3LE state of **DPTZ-DTO** is 2.67 eV), thus TADF is expected, and indeed it was observed. In polar solvents, the $^3CS/^3LE$ states energy gap becomes larger, and no TADF was observed. The three compounds in which the PTZ unit is oxidized have a large energy gap between the 3LE state and the CS state, therefore it is unlikely to observe significant TADF. Interestingly, the CS state energy varied noticeably with one-atom alternation (Se vs. S), which may impose substantial effects on the TADF property. Therefore, one should be careful in predicting heavy atom effects in TADF emitters.⁸⁰

$$E_{CS} = e[E_{OX} - E_{RED}] + \Delta G_S \quad (2)$$

$$\Delta G_{CS} = e[E_{OX} - E_{RED}] - E_{00} + \Delta G_S \quad (3)$$

$$\Delta G_S = -\frac{e^2}{4\pi\epsilon_S\epsilon_0 R_{CC}} - \frac{e^2}{8\pi\epsilon_0} \left(\frac{1}{R_D} + \frac{1}{R_A} \right) \left(\frac{1}{\epsilon_{REF}} - \frac{1}{\epsilon_S} \right) \quad (4)$$

Femtosecond transient absorption spectra

We analyzed the photoinduced behavior of the compounds by measuring their transient absorption spectra with 150 fs time resolution (Fig. 5). All compounds were excited at 350 nm, and transient data were acquired in a time interval spanning up to 3.6 ns. The kinetic traces were then fitted using a global analysis procedure and applying a linear unidirectional decay scheme. The fitting procedure allowed us to extract the time constants describing the excited state evolution and the associated spectral components, termed Evolution Associated Difference Spectra (EADS).^{83,84} The EADS obtained by analyzing the data recorded for the compounds **PTZ-DTO**, **PSeZ-DTO** and **DPTZ-DTO** in TOL solution are presented in Fig. 5.

The spectra of all compounds are qualitatively similar, although some difference can be noticed in the kinetic constants extracted from the fitting. In all cases the initial EADS, living less than 0.5 ps, shows an intense peak centered at about 400 nm. According to previous literature,⁸⁵ this component is attributed to the response of the solvent, specifically to its two-photon absorption, which is promoted upon excitation with UV light. The following spectral component (red EADS in Fig. 5) appears as a very broad positive band, extending within all the probed spectral interval, for all samples.^{86,87}

Based on the transient spectrum of the reference compound **Br-DTO** (Fig. S13), we attribute this component to the S_1 singlet excited state of the compounds. For **PTZ-DTO** this state evolves in about 7.6 ps. The following EADS present two peaks, one centered at about 420 nm and the other at about 520 nm. Based on the results obtained from spectroelectrochemistry, these bands can be respectively assigned to the $DTO^{\cdot-}$ anion and the $PTZ^{\cdot+}$ cation, thus indicating the occurrence of photoinduced CS (Fig. S16).

The intensity of the signal slightly decreases on the following 38.8 ps timescale, and on this same timescale the band at 420 nm experiences a small blue shift. We attribute this

Table 4 Gibbs free energy changes for charge separation (ΔG_{CS}) and charge separated states energy (E_{CS}) of the compounds in different solvents^a

	ΔG_{CS} (eV)					E_{CS} (eV)				
	HEX	TOL	DCM	MeOH	ACN	HEX	TOL	DCM	MeOH	ACN
PTZ-DTO ^b	-0.77	-0.94	-1.44	-1.57	-1.58	+2.90	+2.73	+2.23	+2.10	+2.09
PSeZ-DTO ^c	-0.70	-0.86	-1.31	-1.42	-1.43	+2.95	+2.79	+2.34	+2.23	+2.22
DPTZ-DTO ^d	-0.60	-0.79	-1.32	-1.46	-1.47	+3.03	+2.84	+2.31	+2.17	+2.16
PTZ-O-DTO ^e	+0.01	-0.18	-0.70	-0.84	-0.85	+3.64	+3.46	+2.94	+2.80	+2.79
PTZ-O2-DTO ^f	+0.03	-0.14	-0.64	-0.76	-0.77	+3.73	+3.56	+3.06	+2.93	+2.93
DPTZ-O-DTO ^g	-0.18	-0.35	-0.84	-0.97	-0.98	+3.46	+3.29	+3.80	+2.67	+2.66

^a Cyclic voltammetry was recorded in deaerated solutions containing 0.10 M Bu_4NPF_6 . Pt electrode is the counter electrode, glassy carbon electrode is the working electrode, and $Ag/AgNO_3$ couple is the reference electrode. ^b $E_{00} = 3.67$ eV. ^c $E_{00} = 3.65$ eV. ^d $E_{00} = 3.63$ eV. ^e $E_{00} = 3.64$ eV. ^f $E_{00} = 3.70$ eV. ^g $E_{00} = 3.64$ eV. E_{00} ($E_{00} = 1240/\lambda$) is the singlet state energy of the compounds, λ is the wavelength of the crossing point between the normalized UV/vis absorption spectrum and the fluorescence spectrum.





Fig. 5 EADS obtained from global analysis of the transient absorption data recorded for (a) PTZ-DTO, (b) PSeZ-DTO, (c) DPTZ-DTO, (d) PTZ-O-DTO, (e) PTZ-O2-DTO and (f) DPTZ-O-DTO in TOL upon excitation at 350 nm. Panels (g)–(i) report a comparison of the kinetic traces recorded at the maximum of the PTZ⁺ signal for the three compounds presented in (a)–(c) in TOL and benzonitrile (BZN). The compounds were excited at 350 nm.

evolution to a stabilization of the CS state induced by solvent rearrangement and vibrational cooling. The lifetime of the CS state is longer than the timescale accessed with our setup, indicating that charge recombination (CR) is expected to occur on a timescale longer than 5 ns.

Photoinduced CS was also observed for both **PSeZ-DTO** (Fig. 5b) and **DPTZ-DTO** (Fig. 5c). For the compound containing Se, the dynamics of CS appears a bit slower, in agreement with the modified oxidation potential of the electron donor occurring upon substituting S with Se. For this sample we observe a fast 1 ps component, attributed to a solvent induced relaxation of the S₁ state, followed by CS, which occurs within 15 ps. CR is not observed within the investigated timescale also in case of **PSeZ-DTO**. The dynamics of **DPTZ-DTO** (Fig. 5c) is very similar to that of **PTZ-DTO**. For this sample, CS occurs in 9.9 ps and CR is expected to occur in >5 ns.

Fs-TA measurements were then repeated in a more polar solvent, benzonitrile (BZN) (Fig. S14), where the samples are

quite soluble. In all cases we observed the occurrence of photoinduced CS. In the polar medium the lifetime of the CS state is shortened, as expected because of its stabilization in a more polar environment. Although a complete CR is still not observed, the cation and anion bands significantly decrease in intensity within the investigated time interval, suggesting an increased CR rate, as shown from the comparison of kinetic traces reported in Fig. 5.

Finally, we measured the transient spectra of the oxidized compounds **PTZ-O-DTO**, **PTZ-O2-DTO** and **DPTZ-O-DTO** in TOL, whose EADS are also reported in Fig. 5. The experimental results showed that narrower peaks were observed at 420 nm for these three compounds, while broader absorption bands were observed at 550 nm, and the spectral shapes were slightly different from those of the unoxidized compounds. For both **PTZ-O-DTO** and **PTZ-O2-DTO**, the final EADS presents an increased intensity in the short wavelength region (<400 nm), indicating the transition towards the localized triplet state,



which absorbs at 350 nm, as evidenced by measurements performed on a longer timescale (see next paragraph). Measurements were repeated in BZN also for these samples. The transient absorption spectral characteristics of the oxidized compounds are similar to those observed in TOL, but there are some differences in the time constants of the kinetic processes (Fig. S15).

Nanosecond transient absorption spectra

We then checked the slower dynamics by measuring the ns-TA spectra of the compounds in deaerated TOL. For **PTZ-DTO**, we observed a broad positive absorption band centered at 465 nm and extending in the 400–600 nm spectral range (Fig. 6a). Based on the spectroelectrochemical results (Fig. S16), this band is assigned to a mixture of ^3LE and ^3CS states. This result is consistent with the DFT theoretical calculations. The lifetime of the transient species was determined as 1.6 μs (Fig. 6d), which is close to the delayed fluorescence lifetime of this compound (1.4 μs). Moreover, the lifetimes determined by monitoring the decay traces at 350 nm and 465 nm are same, indicating that the ^3LE and ^3CS states are in equilibrium. This observation is different from a previous study, showing only the ^3LE state for some TADF emitters when using ns-TA spectroscopy.⁴²

In case of **PSeZ-DTO** (Fig. 6b), we noticed two distinct absorption peaks, centered at 420 nm and 510 nm. Based on the spectroelectrochemistry outcomes (Fig. S16), we assign the band centered at 420 nm to the $\text{DTO}^{\cdot-}$ radical anion, and that at 510 nm to the $\text{PSeZ}^{\cdot+}$ radical cation. Another absorption band

was also observed, centered at 350 nm, which is assigned to the ^3LE state. The lifetime of **PSeZ-DTO** was determined as 2.2 μs (Fig. 6b). Interestingly, the transient lifetime of **PSeZ-DTO** resulted longer than that of **PTZ-DTO**, which is against the previous reports, claiming that RISC can be enhanced by incorporation of Se atom in the molecule.^{58–61} For **DPTZ-DTO** (Fig. 6c), we also observed an absorption band centered at 420 nm, attributed to the anion, and the band centered at 510 nm, attributed to the cation, besides an absorption band centered at 350 nm attributed to $^3\text{DTO}^*$ (Fig. S18). The excited state lifetime was determined as 0.8 μs , which is consistent with the delayed fluorescence lifetime (0.8 μs). Thus, for all compounds we observed both the ^3LE and ^3CS states in the ns-TA spectrum, which provides an experimental basis for the spin-vibronic coupling mechanism of TADF.^{25,50}

We also measured the ns-TA spectra of the three compounds in HEX (Fig. S17). For **PTZ-DTO**, we observed a positive absorption band centered at 465 nm, narrower than that observed in TOL. For **PSeZ-DTO**, the absorption band centered at 350 nm, which is attributed to the ^3LE state, is more significant. This result demonstrates that in HEX the CS state of **PSeZ-DTO** has higher energy than the ^3LE state, and that the ^3CS state is not significantly populated, also suppressing TADF. For **DPTZ-DTO**, we still observed a broad positive band centered at 465 nm and ranging between 400 and 600 nm, assigned to the mixture of ^3LE and ^3CS states. For this compound, the ^3LE and ^3CS state energy levels are close to each other both in HEX and TOL solutions, making TADF efficient in both solvents (Fig. S17). Finally, we measured the ns-TA spectra of **PTZ-O**



Fig. 6 Nanosecond transient absorption spectra of (a) **PTZ-DTO**, (b) **PSeZ-DTO** and (c) **DPTZ-DTO**. Decay curves of (d) **PTZ-DTO**, (e) **PSeZ-DTO** at 460 nm and (f) **DPTZ-DTO** at 530 nm after pulsed laser excitation at 355 nm. $c = 2.0 \times 10^{-5}$ M in deaerated TOL, 25 °C.



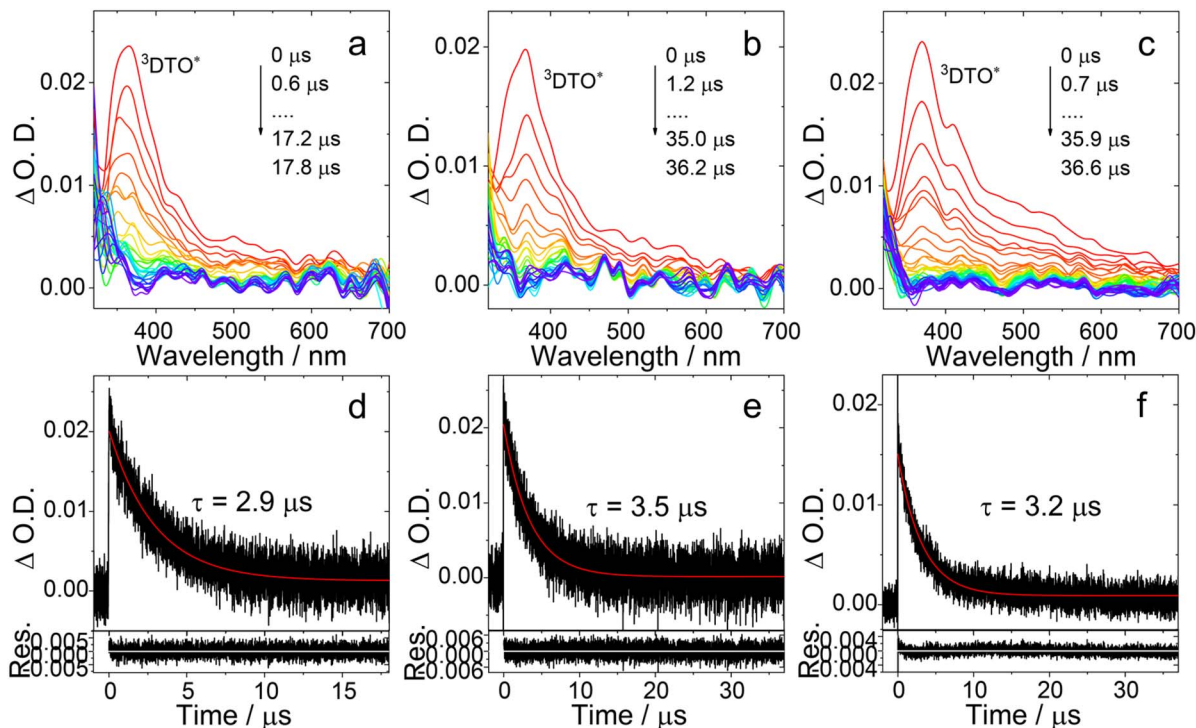


Fig. 7 Nanosecond transient absorption spectra of (a) PTZ-O-DTO, (b) PTZ-O₂-DTO and (c) DPTZ-O-DTO. Decay traces of (d) PTZ-O-DTO, (e) PTZ-O₂-DTO and (f) DPTZ-O-DTO at 370 nm after pulsed laser excitation at 355 nm. $c = 2.0 \times 10^{-5}$ M in deaerated TOL, 25 °C.

DTO, PTZ-O₂-DTO and DPTZ-O-DTO (Fig. 7), observing similar spectra. For all of them we notice an absorption band centered at 360 nm, attributed to the ³LE state, while the ³CS state is no longer observed. These results infer that the energy of the ³CS state is higher than that of the ³LE state, in agreement with the electrochemical studies and DFT calculation (Table 4). Moreover, we did not observe the rise phase of the ns-TA spectra of the compounds, indicating that the formation of the long-lived species (³CS state and the ³LE state) is fast and occurs on a shorter time scale than the instrument response function (IRF) of the ns-TA spectrometer (*ca.* 10 ns).

Based on the ns-TA spectral data and the delayed fluorescence quantum yields of the compounds, we calculated the RISC rate constants according to a literature method.⁴¹ For PTZ-DTO, the rate constants of the RISC was calculated as $k_{\text{RISC}} = 6.1 \times 10^6 \text{ s}^{-1}$, for PSeZ-DTO it resulted $k_{\text{RISC}} = 4.9 \times 10^6 \text{ s}^{-1}$, and a value of $k_{\text{RISC}} = 3.6 \times 10^6 \text{ s}^{-1}$ was obtained for DPTZ-DTO. It is clear that, at least in our case, incorporation of heavy atom Se in the molecular structure does not enhance the reverse ISC process, although in some Se-substituted TADF emitters⁵⁸ an heavy atom effect on the RISC process was indeed observed (with replace of S atom with Se atom, the k_{RISC} of the TADF emitters increased from $1.2 \times 10^4 \text{ s}^{-1}$ to $k_{\text{RISC}} = 2.8 \times 10^5 \text{ s}^{-1}$). For the representative TADF compound 4Cz-IPN,⁸⁴ k_{RISC} resulted $6.8 \times 10^6 \text{ s}^{-1}$, while for the brominated or iodinated 4Cz-IPN, the RISC rate constant was $k_{\text{RISC}} = 7 \times 10^7 \text{ s}^{-1}$. Therefore, in this case the so-called heavy atom effect for the RISC was observed. However, it should be pointed out that

other factors, such as the planarity of the chromophores, also play a role in determining the RISC rate constants.^{35,41,85} Thus, one should be careful in the interpretation of the photo-physical properties of the electron donor-acceptor TADF emitters.

Time-resolved electron paramagnetic resonance (TREPR) spectra

We then investigated the transient paramagnetic species involved in the TADF process using pulsed laser excited TREPR (Fig. 8). In the absence of an external magnetic field, the splitting of triplets is described by the Zero Field Splitting (ZFS) D and E parameters, the main contribution to which is given by the dipole-dipole interaction between the two unpaired electrons of the triplet state. The ³CS state and the ³LE state can be easily discriminated with this technique, based on their distinctly different ZFS D parameters.^{22,48,49,74} ³LE state usually have a larger ZFS D parameter than the ³CS state because of the smaller average distance between the electrons for the ³LE state. It is known that during photoexcitation of chromophores, the sublevels of the photoexcited triplet states are populated selectively, *i.e.* an electron spin polarization (ESP) is resulted and the ESP phase pattern of the TREPR spectra differs depending on which sublevels are overpopulated. Moreover, the ISC mechanism leading to the population of the ³LE state can be also elucidated, based on the ESP of the triplet excited state TREPR spectra,^{50,88-92} although only the RP-ISC mechanism gives a unique ESP phase pattern of the triplet state.²²





Fig. 8 Experimental (black) and simulated (red) TREPR spectra of (a) PTZ-DTO, (b) PSeZ-DTO, (c) DPTZ-DTO, (d) PTZ-O-DTO, (e) PTZ-O₂-DTO and (f) DPTZ-O-DTO with parameters given in Table 5. The experimental spectra are shown for a 500 ns delay after the laser flash with an integration window of 100 ns. The laser excitation wavelength is 355 nm, $c = 5.0 \times 10^{-4}$ M in solvents of TOL/2MeTHF (3/1, v/v). Blue and purple color represent the decomposition of the experimental spectrum into two spectra with similar D , E parameters but multiplet and net polarizations, respectively. Three different triplet states were used in the simulation spectrum of DPTZ-DTO (see Table 5 for details). The spin-polarised patterns are characterised by absorptive (A) and emissive (E) features: positive features are in enhanced absorption (A) and negative features are in emission (E).

The spectra of **PTZ-DTO** were acquired in a frozen solution with mixed solvent TOL/2-methylTHF (3 : 1, v/v) (Fig. 8a). The obtained spectrum is typical for a triplet state of powder sample (randomly oriented molecules in the magnetic field), however, the ESP phase pattern, which is (e, a, e, a, e, a) (e: emission; a: enhanced absorption), corresponds to the overpopulation of the T_y sublevels of the triplet state and is different from that of an ordinary triplet state formed by the spin-orbit coupling (SOC) ISC mechanism, which usually give an ESP phase pattern (e, e, e, a, a, a) or (a, a, a, e, e, e). This result allows us to assume that the triplet state of **PTZ-DTO** may be produced by a mechanism other than SOC-ISC, as for instance, the SOCT-ISC,⁴³ considering the orthogonal geometry of the molecular structures.⁹³ A RP-ISC mechanism can be excluded based on the ESP of the triplet state.^{48,94,95} The ZFS of $|D|$ and $|E|$ were determined as 3400 MHz (121.3 mT) and 425 MHz (15.2 mT), respectively, by simulation of the spectrum (the sign of D is not defined, but E and D values are of the opposite sign, Table 5).

We also measured the TREPR spectrum of **PTZ-DTO** in a solvent with slightly higher polarity (Fig. S19), observing a different polarization pattern. The spectrum is described as

the sum of the spectra of the localized 3LE state, with a ZFS D parameter of 2900 MHz and the spectrum of the SCRPs. To describe the experimentally observed asymmetry of the triplet spectrum, we added a signal from the net polarization of the 3LE state (second component), although the asymmetry may have a different nature. A satisfactory simulation of the spectrum requires three components (Table S1), two species are attributed to the 3LE state, with a ZFS D parameter of 2900 MHz and with different polarization types: multiplet and net polarizations. The third component is assigned as a SCRPs from a triplet precursor,⁹⁸ the simulation gives a D_z magnitude of -200 MHz assuming dipole interaction in a SCRPs, but the contribution from the exchange interaction cannot be excluded. The CS state in the TREPR of **PTZ-DTO** in DCM/2MeTHF is in agreement with observations made in the fs-TA spectra.

For **PSeZ-DTO** (Fig. 8b), the ESP phase pattern of the triplet state TREPR spectrum is (a, a, a, e, e, e), which is typical for a triplet state formed by the normal SOC-ISC.^{88,94} The spectral simulation indicates that the ZFS D parameter is 2900 MHz, and the population ratio is $p_x : p_y : p_z = 0.00 : 0.00 : 1.00$. A similar narrow signal was also observed for the **DPTZ-DTO** compound, in



Table 5 Zero field splitting parameters (D and E) and relative population ratios p_x , p_y , p_z of the zero field spin states of the compounds^a

Compounds	$ D $ (MHz); $ E $ (MHz) ^b	$p_x : p_y : p_z$
PTZ-DTO	[3400; 425]	0.00 : 1.00 : 0.00
PSeZ-DTO	[2900; 480], spec. 1	0.00 : 0.00 : 1.00
	[2900; 480], spec. 2	Net-polarization
DPTZ-DTO	[2700; 450], spec. 1	0.00 : 0.17 : 0.83
	[3400; 425], spec. 2	0.00 : 1.00 : 0.00
PTZ-O-DTO	SCRIP, Dz = -200, spec. 3	Triplet precursor
	[3400; 142], spec. 1	0.26 : 0.00 : 0.74
	[3400; 142], spec. 2	Net-polarization
PTZ-O2-DTO	[3400; 162], spec. 1	0.31 : 0.00 : 0.69
	[3400; 162], spec. 2	Net-polarization
DPTZ-O-DTO	[3400; 162], spec. 1	0.33 : 0.00 : 0.67
	[3400; 162], spec. 2	Net-polarization

^a Obtained from simulations of the triplet state TREPR spectra of the indicated compounds in TOL/2MeTHF (1/3) at 80 K. g -Factor is 2.0030 ± 0.0005 for all spectra. Gaussian line width (peak-to-peak) is 20 mT for spectra with net polarization and it is 2 mT for all other spectra. The error of the ZFS parameters does not exceed 5%. ^b It is impossible to determine the sign of the parameters from the analysis of the EPR spectra, therefore the absolute values of the parameter D and the sign of E are given from the condition $(-1/3D \leq E \leq 0)$.^{96,97}

contrast to the **PTZ-DTO** compound. A narrow signal was observed for **DPTZ-DTO** both for the frozen of DCM/2MeTHF and TOL/2MeTHF. The spectrum of the triad **DPTZ-DTO** is more complicated than that of **PTZ-DTO** (Fig. 8c). A satisfactory simulation requires three components, two of them, having the ZFS D parameters of 2700 MHz and 3400 MHz respectively, are assigned as two different triplet states, showing different population ratios of three sublevels. The spectrum with D parameter of 3400 MHz completely coincides with spectrum of **PTZ-DTO** (Table 5). Increasing the delocalization in this case leads to a decrease in

the parameter D . The third component is as already noted a SCRIP, with electron spin dipolar interaction of -200 MHz (Table 5), and it appears as the precursor of the T_0 substrate. Since **DPTZ-DTO** shows significant TADF, we propose that this is another example of interconversion between the 3LE state and the 3CS state, *i.e.*, this is again an experimental evidence for the spin-vibronic coupling model of the TADF mechanism.

The effect of oxidation of the PTZ moiety on the photo-physics of the TADF emitters was also studied with TREPR spectroscopy (Fig. 8d). The ESP of the triplet state TREPR spectrum of **PTZ-O-DTO** is (a, a, a, e, e, e), which is the typical ESP phase pattern of a triplet state produced by SOC-ISC.^{88,94} The spectral simulation gives a ZFS D parameter of 3400 MHz, which is literally the same of that of **PTZ-DTO** (Table 5). However, the population ratio of the triplet state is $p_x : p_y : p_z = 0.26 : 0.00 : 0.74$, which is different from that of **PTZ-DTO**.

These results indicate that there is an equilibrium between 3LE state and 3CS state, supporting the spin-vibronic mechanism for TADF. Other TADF molecules have been investigated by TREPR spectroscopy in previous studies, showing a 3CS state with a D value from 430 to 1440 MHz.^{43,49,51,67}

Thus, we assume that in **PTZ-DTO**, the ISC mechanism is probably SOCT-ISC, whereas for **PTZ-O-DTO**, it is SOC-ISC. Note that the triplet spectrum of **PTZ-O-DTO** has no symmetric form which may be a result of the triplet net polarization. Considering the contribution from net polarization of this triplet gives a satisfactory description of the experimental spectrum (Table 5). Further evidence for this conclusion comes from **PTZ-O2-DTO** (Fig. 8e). For this compound, a TREPR spectrum with (a, a, a, e, e, e) electron spin polarization phase pattern was observed, which is attributed to a 3LE state. The simulation of the spectrum gives a ZFS D parameter of 3400 MHz, similar to that of **PTZ-O-DTO** ($D = 3400$ MHz) and **PTZ-DTO** ($D = 3400$ MHz) (Table 5).



Fig. 9 Optimized ground state geometry of (a) PTZ-DTO, (b) PSeZ-DTO, (c) DPTZ-DTO, (d) PTZ-O-DTO, (e) PTZ-O2-DTO and (f) DPTZ-O-DTO in gas phase, the blue and pink sheets show the planes of the donor and the acceptor, respectively, isovalue = 0.02. The selected dihedral angles between the electron donor and acceptor are presented. Calculations were performed with DFT theory at the B3LYP/6-31G(d) level with Gaussian 16.



Similar results were obtained for **DPTZ-O-DTO** (Fig. 8f and Table 5). Thus, we propose that for all the oxidized compounds, the T_1 state (3LE) is confined on the DTO unit. The above TREPR spectral studies demonstrated a few critical aspects of the studied electron D–A compounds. First, the ISC mechanism of these TADF emitters most likely is SOCT-ISC, not the normal SOC-ISC. Second, the CS state (long-lived one) is formed *via* the 3LE state, thus verifying the spin-vibronic coupling model. Theoretical computations can reproduce the ZFS D magnitudes to some extents (Table S2). We also calculated the ZFS principal axis (Fig. S26). However, we did not make a detailed correlation between the theoretical calculated population ratios of the three sublevels of the T_1 state of the compounds with the TREPR spectral results.

Theory calculations

Finally, we performed a theoretical analysis to substantiate and further interpret our data. The ground state geometry of the compounds has been optimized at DFT-B3LYP/6-31G(d) level (Fig. 9). The dihedral angle between DTO and PTZ resulted *ca.* 80°, which is close to a vertical configuration. PTZ presents a puckered conformation at the ground state, while DTO adopts a planar geometry. The triplet state spin density of the compounds with PTZ unit oxidized in TOL were also analyzed (Fig. 10). For **PTZ-DTO**, **PSeZ-DTO** and **DPTZ-DTO**, the triplet state density is distributed on PTZ and DTO. The ns-TA spectra indicated that the T_1 state is not a pure $^3DTO^*$ state or a pure 3CS state, but an equilibrium between 3LE / 3CS is established. For the oxidized compounds **PTZ-O-DTO** and **PTZ-O2-DTO** (Fig. 10), the triplet state density is confined on the DTO moiety, confirming that the T_1 state is a 3LE state, as experimentally observed. Other functional and basis sets were tested, but generally the same results were observed (Fig. S25).

In the polar solvent ACN, the spin density surface of the T_1 state of the dyads with the PTZ part oxidized is localized on the DTO moieties (Fig. S24). These results are in agreement with ns-TA spectroscopy (Fig. 7). For the dyads and the triad where the PTZ unit is oxidized, the T_1 state is also a 3LE state, in agreement with the TREPR results. For **DPTZ-O-DTO** the triplet state spin density is shown to be a mixture of 3LE and 3CS states, whereas only 3LE state is observed in ns-TA and TREPR spectra (Fig. 7 and 8).



Fig. 10 Spin density surfaces of the dyads in the T_1 state of (a) **PTZ-DTO**, (b) **PSeZ-DTO**, (c) **DPTZ-DTO**, (d) **PTZ-O-DTO**, (e) **PTZ-O2-DTO** and (f) **DPTZ-O-DTO** in TOL, isovalue = 0.02. Calculations were performed with DFT at the UB3LYP/6-31G(d) level with Gaussian 16.



Scheme 2 Photophysical process of (a) **PTZ-DTO** and (b) **PTZ-O-DTO** in TOL. The 1LE and 3CS state energy is calculated by TDDFT at the B3LYP/6-31G(d) level using Gaussian 16. The 1CS state energy is obtained by electrochemical calculations (the energy levels of the 1CS state obtained from TDDFT are as follows: **PTZ-DTO**: 2.62 eV; **PTZ-O-DTO**: 3.42 eV) the 3LE state energy is estimated by the vibrational $00 T_1 \rightarrow S_0$ transition based on phosphorescence spectra.

The photophysical processes of **PTZ-DTO** and **PTZ-O-DTO** are summarized in Scheme 2. Upon photoexcitation, the 1LE state is populated for both compounds, which undergoes a rapid charge separation to give the 1CS state. For **PTZ-DTO**, the 1CS , 3CS and 3LE states have similar energies, and the 3LE and 3CS states are observed simultaneously in the ns-TA spectra, justifying the observed TADF. For the compounds with the oxidized PTZ unit, the energy of the CS state increases, while no change in energy occurs for the 3LE state energy. These compounds thus present a large energy gap between the 3LE state and the CS state, and consequently only the 3LE state is observed in the ns-TA spectra. In this case the spin-vibronic coupling between 3LE and 3CS is weak and it is unlikely to observe significant TADF.

Conclusions

In summary, we studied the electronic excited states dynamics and electron spin dynamics of a series representative thermally activated delayed fluorescence (TADF) emitters based on phenothiazine–dibenzothiophene-*S,S*-dioxide (**PTZ-DTO**) dyads using steady state and femtosecond/nanosecond transient absorption (fs/ns-TA) spectroscopy, as well as pulsed laser-excited time-resolved electron paramagnetic resonance (TREPR) spectroscopy. The molecular structures were systematically varied by attaching an extra electron donor (PTZ) to the electron acceptor (DTO), using Se atom to replace the S atom or by



oxidation of the PTZ unit to sulfoxide and sulfone. Fs-TA spectroscopy shows that charge separation (CS) occurs readily (always in <20 ps) in all the dyads where the donor is not oxidized, while charge recombination (CR) is quite slow (>5 ns) in toluene, but is accelerated in polar solvents. Ns-TA spectroscopy demonstrated the coexistence of the ³CS and ³LE states for the non-oxidized dyads. Oxidation of the PTZ unit increases the energy of the CS state and as a result, only the ³LE (localized on the DTO moiety) state is observed. In the TREPR spectra we observed an electron spin polarization (ESP) phase pattern (e, a, e, a, e, a) for the **PTZ-DTO** dyad, suggesting triplet formation through the spin-orbit charge transfer intersystem crossing (SOCT-ISC) mechanism, rather than the spin-orbit coupling ISC (SOC-ISC) mechanism. However, upon Se substitution, SOC-ISC becomes dominant. In solvents with slightly higher polarity, a spin-correlated radical pair (SCRPA) was also observed (dipolar interaction is -200 MHz), supporting the spin-vibronic coupling mechanism for TADF. These studies are useful for an in-depth understanding of the photophysics of the electron donor-acceptor TADF emitters and for the design of new light emitting materials for OLED applications. A key finding is that the strategy proposed in previous studies, introducing heavy atoms to reduce the efficiency roll-off in OLED devices, is not a universal rule. This study found that in our compounds the introduction of heavy atoms does not perturb the decay lifetimes of delayed fluorescence and triplet states. This result has significant implication for molecular structure design of electron donor-acceptor dyads as OLED emitters. When designing TADF molecules, the impact of structural modifications on the CS state energy level should also be evaluated. Moreover, deep understanding the factors dictating the triplet state lifetime of TADF emitters is important for device performance, because a long-lived triplet state may lead to efficiency roll-off of the OLED devices.

Author contributions

Y. P. and A. A. S. contributed equally to this work. J. Z. conceived the work, acquired the funding, directed the data analysis and the manuscript writing; Y. P. synthesized the compounds, measured and analyzed steady-state optical spectra, electrochemistry, nanosecond transient absorption spectra, performed parts of the data analysis, and modified the manuscript; A. A. S. and V. K. V. measured the TREPR spectra and analyzed the data; G. S., L. B. and M. D. D. did the femtosecond transient absorption studies and analyzed the data. X. C., Y. L. and Y. H. revised parts of the manuscript; X. L. and H. G. performed the theoretical calculations and analyzed the data; all authors contributed to the writing of the manuscript.

Conflicts of interest

There are no conflicts to declare.

Data availability

All relevant data are presented in the main text and/or SI. Supplementary information: the molecular structural

characterization data, UV-Vis absorption and fluorescence spectra, electrochemical studies, nanosecond and femtosecond transient absorption spectra, TREPR spectra and theoretical computation information. See DOI: <https://doi.org/10.1039/d5sc03644e>.

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Notes and references

- 1 H. Tanaka, K. Shizu, H. Miyazaki and C. Adachi, *Chem. Commun.*, 2012, **48**, 11392–11394.
- 2 H. Uoyama, K. Goushi, K. Shizu, H. Nomura and C. Adachi, *Nature*, 2012, **492**, 234–238.
- 3 J. Zhou, P. Chen, X. Wang, Y. Wang, Y. Wang, F. Li, M. Yang, Y. Huang, J. Yu and Z. Lu, *Chem. Commun.*, 2014, **50**, 7586–7589.
- 4 Z. Yang, Z. Mao, Z. Xie, Y. Zhang, S. Liu, J. Zhao, J. Xu, Z. Chi and M. P. Aldred, *Chem. Soc. Rev.*, 2017, **46**, 915–1016.
- 5 X. Cao, D. Zhang, S. Zhang, Y. Tao and W. Huang, *J. Mater. Chem. C*, 2017, **5**, 7699–7714.
- 6 Y. Im, S. Y. Byun, J. H. Kim, D. R. Lee, C. S. Oh, K. S. Yook and J. Y. Lee, *Adv. Funct. Mater.*, 2017, **27**, 1603007.
- 7 T. Huang, W. Jiang and L. Duan, *J. Mater. Chem. C*, 2018, **6**, 5577–5596.
- 8 X. Cai and S.-J. Su, *Adv. Funct. Mater.*, 2018, **28**, 1802558.
- 9 G. C. Hoover and D. S. Seferos, *Chem. Sci.*, 2019, **10**, 9182–9188.
- 10 F. Ni, N. Li, L. Zhan and C. Yang, *Adv. Opt. Mater.*, 2020, **8**, 1902187.
- 11 L. Chen, W.-C. Chen, Z. Yang, J.-H. Tan, S. Ji, H.-L. Zhang, Y. Huo and C.-S. Lee, *J. Mater. Chem. C*, 2021, **9**, 17233–17264.
- 12 Y. Liu, X. Xiao, Y. Ran, Z. Bin and J. You, *Chem. Sci.*, 2021, **12**, 9408–9412.



- 13 B. Xu, Z. Song, M. Zhang, Q. Zhang, L. Jiang, C. Xu, L. Zhong, C. Su, Q. Ban, C. Liu, F. Sun, Y. Zhang, Z. Chi, Z. Zhao and G. Shi, *Chem. Sci.*, 2021, **12**, 15556–15562.
- 14 Y.-Z. Shi, H. Wu, K. Wang, J. Yu, X.-M. Ou and X.-H. Zhang, *Chem. Sci.*, 2022, **13**, 3625–3651.
- 15 T. Zhang, Y. Xiao, H. Wang, S. Kong, R. Huang, V. Ka-Man Au, T. Yu and W. Huang, *Angew. Chem., Int. Ed.*, 2023, **62**, e202301896.
- 16 J. M. D. Santos, D. Hall, B. Basumatary, M. Bryden, D. Chen, P. Choudhary, T. Comerford, E. Crovini, A. Danos, J. De, S. Diesing, M. Fatahi, M. Griffin, A. K. Gupta, H. Hafeez, L. Hammerling, E. Hanover, J. Haug, T. Heil, D. Karthik, S. Kumar, O. Lee, H. Li, F. Lucas, C. F. R. Mackenzie, A. Mariko, T. Matulaitis, F. Millward, Y. Olivier, Q. Qi, I. D W Samuel, N. Sharma, C. Si, L. Spiering, P. Sudhakar, D. Sun, E. T. Te, M. D. Tonet, J. Wang, T. Wang, S. Wu, Y. Xu, L. Zhang and E. Zveman-Colman, *Chem. Rev.*, 2024, **124**, 13736–14110.
- 17 H. Tanaka, K. Shizu, H. Nakanotani and C. Adachi, *J. Phys. Chem. C*, 2014, **118**, 15985–15994.
- 18 T. J. Penfold, *J. Phys. Chem. C*, 2015, **119**, 13535–13544.
- 19 Q. Peng, D. Fan, R. Duan, Y. Yi, Y. Niu, D. Wang and Z. Shuai, *J. Phys. Chem. C*, 2017, **121**, 13448–13456.
- 20 J. R. Lakowicz and B. R. Masters, *J. Biomed. Opt.*, 2008, **13**, 029901.
- 21 F. D. Maiolo, D. K. A. P. Huu, D. Giavazzi, A. Landi, O. Racchi and A. Painelli, *Chem. Sci.*, 2024, **15**, 5434–5450.
- 22 Z. E. X. Dance, S. M. Mickley, T. M. Wilson, A. B. Ricks, A. M. Scott, M. A. Ratner and M. R. Wasielewski, *J. Phys. Chem. A*, 2008, **112**, 4194–4201.
- 23 P. K. Samanta, D. Kim, V. Coropceanu and J.-L. Brédas, *J. Am. Chem. Soc.*, 2017, **139**, 4042–4051.
- 24 X. Zhang, X. Liu, M. Taddei, L. Bussotti, I. Kurganskii, M. Li, X. Jiang, L. Xing, S. Ji, Y. Huo, J. Zhao, M. Di Donato, Y. Wan, Z. Zhao and M. V. Fedin, *Chem.–Eur. J.*, 2022, **28**, e202200510.
- 25 X. Zhang, X. Zhao, K. Ye and J. Zhao, *Chem.–Eur. J.*, 2023, **29**, e202203737.
- 26 L. Cao, X. Liu, X. Zhang, J. Zhao, F. Yu and Y. Wan, *Beilstein J. Org. Chem.*, 2023, **19**, 1028–1046.
- 27 J. Tang, X. Liu, X. Zhang, J. Zhao and Y. Wan, *New J. Chem.*, 2023, **47**, 22418–22429.
- 28 L. Wang, Q. Ou, Q. Peng and Z. Shuai, *J. Phys. Chem. A*, 2021, **125**, 1468–1475.
- 29 J. Gibson, A. P. Monkman and T. J. Penfold, *ChemPhysChem*, 2016, **17**, 2956–2961.
- 30 M. K. Etherington, J. Gibson, H. F. Higginbotham, T. J. Penfold and A. P. Monkman, *Nat. Commun.*, 2016, **7**, 13680.
- 31 F. B. Dias, T. J. Penfold and A. P. Monkman, *Methods Appl. Fluoresc.*, 2017, **5**, 012001.
- 32 I. Kim, S. O. Jeon, D. Jeong, H. Choi, W.-J. Son, D. Kim, Y. M. Rhee and H. S. Lee, *J. Chem. Theory Comput.*, 2020, **16**, 621–632.
- 33 L. Lv, K. Yuan, T. Zhao and Y. Wang, *J. Mater. Chem. C*, 2020, **8**, 10369–10381.
- 34 X.-K. Chen, *CCS Chem.*, 2020, **2**, 1256–1267.
- 35 M. Hempe, N. A. Kukhta, A. Danos, M. A. Fox, A. S. Batsanov, A. P. Monkman and M. R. Bryce, *Chem. Mater.*, 2021, **33**, 3066–3080.
- 36 I. Kim, K. H. Cho, S. O. Jeon, W.-J. Son, D. Kim, Y. M. Rhee, I. Jang, H. Choi and D. S. Kim, *JACS Au*, 2021, **1**, 987–997.
- 37 S. Lin, Z. Pei, B. Zhang, H. Ma and W. Liang, *J. Phys. Chem. A*, 2022, **126**, 239–248.
- 38 T. Hosokai, H. Matsuzaki, H. Nakanotani, K. Tokumaru, T. Tsutsui, A. Furube, K. Nasu, H. Nomura, M. Yahiro and C. Adachi, *Sci. Adv.*, 2017, **3**, e1603282.
- 39 J. Choi, D.-S. Ahn, K. Y. Oang, D. W. Cho and H. Ihee, *J. Phys. Chem. C*, 2017, **121**, 24317–24323.
- 40 J. Ehrmaier, E. J. Rabe, S. R. Pristash, K. L. Corp, C. W. Schlenker, A. L. Sobolewski and W. Domcke, *J. Phys. Chem. A*, 2019, **123**, 8099–8108.
- 41 R. J. Vázquez, J. H. Yun, A. K. Muthike, M. Howell, H. Kim, I. K. Madu, T. Kim, P. Zimmerman, J. Y. Lee and T. G. Iii, *J. Am. Chem. Soc.*, 2020, **142**, 8074–8079.
- 42 L. Fisher Jr, R. J. Vázquez, M. Howell, A. K. Muthike, M. E. Orr, H. Jiang, B. Dodgen, D. R. Lee, J. Y. Lee, P. Zimmerman and T. Goodson III, *Chem. Mater.*, 2022, **34**, 2161–2175.
- 43 G. Tang, A. A. Sukhanov, J. Zhao, W. Yang, Z. Wang, Q. Liu, V. K. Voronkova, M. Di Donato, D. Escudero and D. Jacquemin, *J. Phys. Chem. C*, 2019, **123**, 30171–30186.
- 44 D. M. E. Freeman, A. J. Musser, J. M. Frost, H. L. Stern, A. K. Forster, K. J. Fallon, A. G. Rapidis, F. Cacialli, I. McCulloch, T. M. Clarke, R. H. Friend and H. Bronstein, *J. Am. Chem. Soc.*, 2017, **139**, 11073–11080.
- 45 X. Xiao, Y. Yan, A. A. Sukhanov, S. Doria, A. Iagatti, L. Bussotti, J. Zhao, M. Di Donato and V. K. Voronkova, *J. Phys. Chem. B*, 2023, **127**, 6982–6998.
- 46 S. Weber, *eMagRes*, 2017, **6**, 255–270.
- 47 M. D. E. Forbes, L. E. Jarocho, S. Sim and V. F. Tarasov, *Adv. Phys. Org. Chem.*, 2013, **47**, 1–83.
- 48 T. Ogiwara, Y. Wakikawa and T. Ikoma, *J. Phys. Chem. A*, 2015, **119**, 3415–3418.
- 49 B. H. Drummond, N. Aizawa, Y. Zhang, W. K. Myers, Y. Xiong, M. W. Cooper, S. Barlow, Q. Gu, L. R. Weiss, A. J. Gillett, D. Credgington, Y.-J. Pu, S. R. Marder and E. W. Evans, *Nat. Commun.*, 2021, **12**, 4532.
- 50 X. Chen, X. Xiao and J. Zhao, *J. Mater. Chem. C*, 2022, **10**, 4546–4557.
- 51 K. Ye, A. A. Sukhanov, Y. Pang, A. Mambetov, M. Li, L. Cao, J. Zhao, V. K. Voronkova, Q. Peng and Y. Wan, *Phys. Chem. Chem. Phys.*, 2025, **27**, 813–823.
- 52 N. Zarrabi, B. J. Bayard, S. Seetharaman, N. Holzer, P. Karr, S. Ciuti, A. Barbon, M. Di Valentin, A. van der Est, F. D'Souza and P. K. Poddutoori, *Phys. Chem. Chem. Phys.*, 2021, **23**, 960–970.
- 53 R. S. Nobuyasu, Z. Ren, G. C. Griffiths, A. S. Batsanov, P. Data, S. Yan, A. P. Monkman, M. R. Bryce and F. B. Dias, *Adv. Opt. Mater.*, 2016, **4**, 597–607.
- 54 F. B. Dias, J. Santos, D. R. Graves, P. Data, R. S. Nobuyasu, M. A. Fox, A. S. Batsanov, T. Palmeira, M. N. Berberan-Santos, M. R. Bryce and A. P. Monkman, *Adv. Sci.*, 2016, **3**, 1600080.



- 55 Z. Yu, A. A. Sukhanov, X. Xiao, A. Iagatti, S. Doria, V. Butera, J. Zhao, V. K. Voronkova, M. Di Donato and G. Mazzone, *J. Phys. Chem. B*, 2023, **127**, 5905–5923.
- 56 S. Gan, W. Luo, B. He, L. Chen, H. Nie, R. Hu, A. Qin, Z. Zhao and B. Z. Tang, *J. Mater. Chem. C*, 2016, **4**, 3705–3708.
- 57 Y. Rout, C. Montanari, E. Pasciucco, R. Misra and B. Carlotti, *J. Am. Chem. Soc.*, 2021, **143**, 9933–9943.
- 58 B. H. Drummond, G. C. Hoover, A. J. Gillett, N. Aizawa, W. K. Myers, B. T. McAllister, S. T. E. Jones, Y.-J. Pu, D. Credgington and D. S. Seferos, *J. Phys. Chem. C*, 2020, **124**, 6364–6370.
- 59 Y.-H. Shi, F. Wang, G.-Y. Sun and Y.-Z. Xie, *Spectrochim. Acta, Part A*, 2022, **264**, 120249.
- 60 W. Shao and J. Kim, *Acc. Chem. Res.*, 2022, **55**, 1573–1585.
- 61 I. S. Park, H. Min and T. Yasuda, *Angew. Chem., Int. Ed.*, 2022, **61**, e202205684.
- 62 J. S. Ward, R. S. Nobuyasu, M. A. Fox, A. S. Batsanov, J. Santos, F. B. Dias and M. R. Bryce, *J. Org. Chem.*, 2018, **83**, 14431–14442.
- 63 R. S. Nobuyasu, J. S. Ward, J. Gibson, B. A. Laidlaw, Z. Ren, P. Data, A. S. Batsanov, T. J. Penfold, M. R. Bryce and F. B. Dias, *J. Mater. Chem. C*, 2019, **7**, 6672–6684.
- 64 M. Hempe, A. K. Harrison, J. S. Ward, A. S. Batsanov, M. A. Fox, F. B. Dias and M. R. Bryce, *J. Org. Chem.*, 2021, **86**, 429–445.
- 65 J. R. Lakowicz, *Principles of Fluorescence Spectroscopy*, Kluwer Academic, New York, 2nd edn, 1999.
- 66 N. J. Turro, V. Ramamurthy and J. C. Scaiano, *Principles of molecular photochemistry: an introduction*, University Science Books, 2009.
- 67 X. Zhao, A. A. Sukhanov, X. Jiang, J. Zhao and V. K. Voronkova, *J. Phys. Chem. Lett.*, 2022, **13**, 2533–2539.
- 68 M. Baruah, W. Qin, C. Flors, J. Hofkens, R. A. L. Vallée, D. Beljonne, M. Van der Auweraer, W. M. De Borggraeve and N. Boens, *J. Phys. Chem. A*, 2006, **110**, 5998–6009.
- 69 S. Sasaki, K. Hattori, K. Igawa and G.-i. Konishi, *J. Phys. Chem. A*, 2015, **119**, 4898–4906.
- 70 T. Nakagawa, S.-Y. Ku, K.-T. Wong and C. Adachi, *Chem. Commun.*, 2012, **48**, 9580–9582.
- 71 S. Y. Lee, T. Yasuda, Y. S. Yang, Q. Zhang and C. Adachi, *Angew. Chem., Int. Ed.*, 2014, **53**, 6402–6406.
- 72 W. Li, J. Li, F. Wang, Z. Gao and S. Zhang, *ACS Appl. Mater. Interfaces*, 2015, **7**, 26206–26216.
- 73 S. Wang, H. Shu, X. Han, X. Wu, H. Tong and L. Wang, *J. Mater. Chem. C*, 2021, **9**, 9907–9913.
- 74 M. K. Etherington, F. Franchello, J. Gibson, T. Northey, J. Santos, J. S. Ward, H. F. Higginbotham, P. Data, A. Kurowska, P. L. Dos Santos, D. R. Graves, A. S. Batsanov, F. B. Dias, M. R. Bryce, T. J. Penfold and A. P. Monkman, *Nat. Commun.*, 2017, **8**, 14987.
- 75 K. Ye, L. Cao, D. M. E. van Raamsdonk, Z. Wang, J. Zhao, D. Escudero and D. Jacquemin, *J. Org. Chem.*, 2022, **18**, 1435–1453.
- 76 R. Ziessel, B. D. Allen, D. B. Rewinska and A. Harriman, *Chem.–Eur. J.*, 2009, **15**, 7382–7393.
- 77 J. Hankache and O. S. Wenger, *Chem.–Eur. J.*, 2012, **18**, 6443–6447.
- 78 M. A. Collini, M. B. Thomas, V. Bandi, P. A. Karr and F. D'Souza, *Chem.–Eur. J.*, 2017, **23**, 4450–4461.
- 79 D. I. Schuster, P. Cheng, P. D. Jarowski, D. M. Guldi, C. Luo, L. Echegoyen, S. Pyo, A. R. Holzwarth, S. E. Braslavsky, R. M. Williams and G. Klichm, *J. Am. Chem. Soc.*, 2004, **126**, 7257–7270.
- 80 G. N. Lim, C. O. Obondi and F. D'Souza, *Angew. Chem., Int. Ed.*, 2016, **55**, 11517–11521.
- 81 V. Mariñas, B. Platzer, J. Labella, F. Caroleo, S. Nardis, R. Paolesse, D. M. Guldi and T. Torres, *Chem.–Eur. J.*, 2022, **28**, e202201552.
- 82 Y. Pei, A. A. Sukhanov, X. Chen, A. Iagatti, S. Doria, X. Dong, J. Zhao, Y. Li, W. Chi, V. K. Voronkova, M. Di Donato and B. Dick, *Chem.–Eur. J.*, 2025, **31**, e202403542.
- 83 I. H. M. van Stokkum, D. S. Larsen and R. van Grondelle, *Biochim. Biophys. Acta, Bioenerg.*, 2004, **1657**, 82–104.
- 84 H. S. Kim, J. Y. Lee, S. Shin, W. Jeong, S. H. Lee, S. Kim, J. Lee, M. C. Suh and S. Yoo, *Adv. Funct. Mater.*, 2021, **31**, 2104646.
- 85 H. Liu, Y. Liu, G. Chen, Y. Meng, H. Peng, J. Miao and C. Yang, *Chem. Sci.*, 2024, **15**, 12598–12605.
- 86 R. Berera, R. van Grondelle and J. T. M. Kennis, *Photosynth. Res.*, 2009, **101**, 105–118.
- 87 X. Xiao, T. Mu, A. A. Sukhanov, Y. Zhou, P. Yu, F. Yu, A. Elmali, J. Zhao, A. Karatay and V. K. Voronkova, *Phys. Chem. Chem. Phys.*, 2023, **25**, 31667–31682.
- 88 H. Levanon and J. R. Norris, *Chem. Rev.*, 1978, **78**, 185–198.
- 89 J. W. Verhoeven, *J. Photochem. Photobiol., C*, 2006, **7**, 40–60.
- 90 T. Biskup, *Front. Chem.*, 2019, **7**, 10.
- 91 M. Imran, X. Zhang, Z. Wang, X. Chen, J. Zhao, A. Barbon and V. K. Voronkova, *Phys. Chem. Chem. Phys.*, 2021, **23**, 15835–15868.
- 92 Z. Wang, X. Zhang and J. Zhao, *J. Phys. Chem. C*, 2021, **125**, 19097–19109.
- 93 Z. Li, W. Shen, X. Liu and R. Liu, *Polym. Chem.*, 2017, **8**, 1579–1588.
- 94 S. Richert, C. E. Tait and C. R. Timmel, *J. Magn. Reson.*, 2017, **280**, 103–116.
- 95 E. W. Evans, Y. Olivier, Y. Puttison, W. K. Myers, T. J. H. Hele, S. M. Menke, T. H. Thomas, D. Credgington, D. Beljonne, R. H. Friend and N. C. Greenham, *J. Phys. Chem. Lett.*, 2018, **9**, 4053–4058.
- 96 C. P. Poole Jr, H. A. Farach and W. K. Jackson, *J. Chem. Phys.*, 1974, **61**, 2220–2221.
- 97 C. E. Tait, M. D. Krzyaniak and S. Stoll, *J. Magn. Reson.*, 2023, **349**, 107410.
- 98 Q. Mi, M. A. Ratner and M. R. Wasielewski, *J. Phys. Chem. A*, 2010, **114**, 162–171.

