

## PAPER

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[View Journal](#) | [View Issue](#)Cite this: *Mater. Adv.*, 2025,  
6, 2691

# A new and facile preparation of 3D urchin-like TiO<sub>2</sub>@graphene core@shell SERS substrates for photocatalytic degradation of RhB†

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Here, we present a new and facile method to grow a graphene (Gr) shell layer on the surface of a 3D urchin-like TiO<sub>2</sub> core layer (TiO<sub>2</sub>@Gr) on a silicon (Si/SiO<sub>2</sub>) substrate as a surface enhanced Raman scattering (SERS) substrate for detecting and degrading rhodamine B (RhB) under UV irradiation. The core layer of 3D urchin-like TiO<sub>2</sub> (UT) was fabricated by a hydrothermal method and a Gr shell layer was grown on the surface of the TiO<sub>2</sub> core layer by a thermal chemical vapor deposition (CVD) method with sodium deoxycholate (SDC) surfactant as a carbon source. The obtained results indicated that the minimum detectable concentration of RhB on the TiO<sub>2</sub>@Gr/SERS substrate-coated graphene (UTG) can reach  $1 \times 10^{-9}$  M with an enhancement factor (EF) of  $1.3 \times 10^5$ . The enhancement of the Raman signals of the UTG can be generated because the graphene acts as the electron acceptor of TiO<sub>2</sub> and prevents charge recombination, and provides indirect charge transfer from TiO<sub>2</sub> to RhB molecules. The UTG SERS substrate can almost completely decompose RhB with a degradation rate of  $0.069 \text{ min}^{-1}$  under UV irradiation at 254 nm within 80 min. The photocatalytic degradation mechanism of UTGs towards RhB was also presented in detail. The results show that the UTG SERS substrate can be further employed for detecting substances and degrading hazardous pollutants.

Received 16th January 2025,  
Accepted 23rd March 2025

DOI: 10.1039/d5ma00040h

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## 1. Introduction

To measure the degree and rate of organic contaminant degradation at low levels, a number of analytical approaches have been developed to monitor the photo-degradation process, including UV-Vis spectrophotometry,<sup>1</sup> gas chromatography-mass

spectrometry (GC-MS),<sup>2</sup> and high-performance liquid chromatography (HPLC).<sup>3</sup> However, UV-Vis spectrophotometry can detect only the concentration of molecules without detecting the chemical structure of the molecules during the degradation of organic contaminants and the intermediates during the photocatalytic process. Compared to UV-Vis, GC-MS and HPLC measurements can be used to identify the intermediates of molecules. Nevertheless, these measurements are still restricted due to quite complicated, costly testing equipment, and are time-consuming.<sup>4</sup> Therefore, it is necessary to develop an *in situ* method to detect the chemical structure of molecules and identify the intermediates during the photocatalytic degradation process of organic contaminants. Recently, the surface-enhanced Raman scattering (SERS) technique has emerged as a powerful tool for detecting and analyzing the structure of chemical and biological molecules due to its high sensitivity, short analysis times, non-destructive determination, easy sample preparation, and lack of requirement for sample extraction and separation.<sup>5,6</sup>

Among semiconductor materials, titanium dioxide (TiO<sub>2</sub>) has been extensively used in the photodegradation of environmental pollutants due to its high photocatalytic activity,

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d5ma00040h>

stability, and nontoxicity.<sup>7–11</sup> However, TiO<sub>2</sub> has weak SERS signals for detecting chemical and biological molecules.<sup>6,7,12</sup> To improve the SERS signals and photocatalytic activity of the TiO<sub>2</sub> material, it is necessary to develop new methods for producing TiO<sub>2</sub> and TiO<sub>2</sub> composite materials with different sizes, shapes, and crystal structures. Compared to zero dimensional (0D), one dimensional (1D) and two dimensional (2D) structures, three dimensional (3D) structures of TiO<sub>2</sub> have advantages such as high surface area-to-volume ratio, low aggregation, and easy separation of particles from solution. Hence, the 3D TiO<sub>2</sub> can provide more space for adsorbing analyte molecules and permit the diffusion of analyte molecules into the structure, and thus will enhance surface reactions.<sup>13</sup> Moreover, to enhance the SERS signals and the photocatalytic degradation, TiO<sub>2</sub> has been combined with nanostructured noble metals (Ag, Au) and/or carbon nanomaterials (graphene, carbon nanotubes).<sup>6–10</sup> With high charge carrier mobility, mechanical flexibility, large surface area, and high absorption of organic molecules through  $\pi$ - $\pi$  interactions, graphene (Gr) is a good candidate for applications in surface-enhanced Raman scattering (SERS)<sup>10,14–16</sup> and photocatalytic degradation.<sup>9,11</sup> Due to its unique optoelectronic and electronic properties, the TiO<sub>2</sub>-Gr hybrid material has emerged as a star material for providing effective charge transfer (CT) between analyte molecules and the substrate surface and electron transport in photocatalysis.<sup>7,10</sup> The TiO<sub>2</sub>-Gr hybrid material has been used to enhance both SERS signals and photocatalytic degradation of environmental pollutants, such as rhodamine 6G,<sup>9</sup> rhodamine B,<sup>17</sup> and tetracycline medicine.<sup>17</sup>

The TiO<sub>2</sub>-Gr hybrid material can be fabricated by some methods, such as thermal chemical vapor deposition (CVD),<sup>18–21</sup> solvo/hydrothermal,<sup>22–24</sup> sol-gel,<sup>25</sup> and mechanical mixing.<sup>26,27</sup> Among these above methods, only the thermal CVD method can synthesize TiO<sub>2</sub>-Gr hybrid materials with a Gr layer covered TiO<sub>2</sub> core to form a TiO<sub>2</sub>@Gr core-shell structure. Some previous studies have used carbon sources such as methane gas (CH<sub>4</sub>),<sup>18,21</sup> acetylene gas (C<sub>2</sub>H<sub>2</sub>),<sup>20</sup> propylene gas (C<sub>3</sub>H<sub>6</sub>), and methanol (CH<sub>3</sub>OH)<sup>19</sup> for the fabrication of TiO<sub>2</sub>@Gr core-shell structure hybrid materials. However, most previous reports have only focused on synthesizing graphene layers directly onto the surface of TiO<sub>2</sub> nanoparticles<sup>19,20</sup> or flat TiO<sub>2</sub> substrates.<sup>18,21</sup> To the best of our knowledge, a 3D urchin-like TiO<sub>2</sub>@Gr core-shell structure hybrid material synthesized by thermal CVD and used as a SERS substrate to enhance the Raman signal and photocatalytic degradation of rhodamine B has not been reported yet.

In the present study, 3D urchin-like TiO<sub>2</sub> (UT) on a silicon (Si/SiO<sub>2</sub>) substrate was fabricated by a hydrothermal method and a Gr shell layer was grown on the surface of the TiO<sub>2</sub> core by the thermal CVD method with sodium deoxycholate (SDC) surfactant as a carbon source. The fabricated 3D urchin-like TiO<sub>2</sub>@Gr core-shell (UTG) samples were used as SERS substrates to detect and decompose rhodamine B (RhB) under UV irradiation. Furthermore, a possible charge transfer mechanism for the SERS enhancement effect and photocatalytic degradation mechanism for UTG towards RhB were also presented in detail.

## 2. Experimental

### 2.1. Materials

Titan tetrachloride (TiCl<sub>4</sub>, 99%), hydrochloric acid (HCl, 37%), absolute ethanol (C<sub>2</sub>H<sub>5</sub>OH, 99.7%), rhodamine B (RhB, 99%), ethylenediaminetetraacetic acid disodium (EDTA, 99%), isopropyl alcohol (IPA, > 99%), and benzoquinone (BQ, 99%) were purchased from Macklin Co. Ltd (China). Gases such as hydrogen (H<sub>2</sub>, > 99%), argon (Ar, > 99%) were purchased from Messer Co. Ltd (Hai Phong, Vietnam). Sodium deoxycholate (SDC, 96%) was purchased from Fujifilm Co. Ltd (Japan).

### 2.2. Fabrication of 3D urchin-like TiO<sub>2</sub>

The 3D urchin-like TiO<sub>2</sub> (UT) was grown on a silicon substrate (Si/SiO<sub>2</sub>) by a hydrothermal method. Firstly, 10 mL HCl was added into 10 mL deionized water, and then 0.3 mL TiCl<sub>4</sub> was added under stirring for 20 min at 7–10 °C. The obtained solution was put into a stainless-steel autoclave of 50 mL volume, followed by placing the Si/SiO<sub>2</sub> substrates (0.5 cm × 0.5 cm) horizontally on the floor of the autoclave. Then, the autoclave was kept at 160 °C for 6 h and then naturally cooled to room temperature. After that, the as-prepared samples were washed with deionized water and ethanol, dried at 50 °C for 30 min and annealed at 400 °C for 1 h in air, and UT samples were obtained.

### 2.3. Fabrication of a 3D urchin-like TiO<sub>2</sub>@Gr core-shell

The 3D urchin-like TiO<sub>2</sub>@Gr (UTG) core-shell was fabricated by the thermal chemical vapor deposition (CVD) method with SDC surfactant as a carbon source for the growth of a graphene (Gr) shell layer. Briefly, a mixed solution of 25  $\mu$ L SDC solution and ethanol with a ratio of 3 : 1 was dropped onto the surface of the TiO<sub>2</sub> samples. Afterwards, the samples were dried for 24 h at room temperature. Before the CVD process for growing a graphene layer, the reactor quartz tube and all gas lines were purged in Ar gas flow of 300 standard cubic centimeter (sccm) for 30 min. Then, the TiO<sub>2</sub> samples covered with SDC were inserted into a quartz tube and heated at 800 °C with a heating rate of 10 °C min<sup>-1</sup> in a mixture of Ar (50 sccm) and H<sub>2</sub> (30 sccm) gases for 30 min. After the CVD process, the H<sub>2</sub> gas was shut off and the samples were cooled down naturally to room temperature in Ar gas flow.

### 2.4. Characterizations

The surface morphologies of the samples were measured by field emission scanning electron microscopy (FESEM, Hitachi S-4800, Japan) at an accelerating voltage of 5 kV, and transmission electron microscopy (TEM, JEM 2100, Joel, Japan) at an accelerating voltage of 200 kV. The crystal structures of samples were determined by X-ray diffraction (XRD, D8 Advance, Bruker, Germany) with Cu-K $\alpha$  radiation ( $\lambda$  = 1.5406 Å) at 40 kV and 20 mA over the 2-theta range of 10°–80°. The energy dispersive X-ray spectrum (EDS) mapping of the samples was determined by SEM (JSM-IT800, Joel, Japan). Photoluminescence (PL) spectra were done by using a spectrofluorometer (FL3-22 JobinYvon Spex, USA).



## 2.5. SERS and photodegradation measurements

SERS and photodegradation properties of UT and UTG core-shell samples were investigated through detecting the Raman vibrational signals of the RhB molecule. 20  $\mu\text{L}$  of RhB solutions with various concentrations ( $10^{-1}$ – $10^{-9}$  M) were uniformly drop coated onto the surface of the samples. The photodegradation performances of the RhB molecules on the samples under ultraviolet light (UV) radiation (UV lamp source with parameters: power (66 W), wavelength (254 nm), length (45 cm), and diameter (2.5 cm)), manufactured by Medicor, Budapest, Hungary, type (BLF-12) were measured at room temperature by a Raman spectrophotometer (LabRAM HR Evolution, Horiba, Japan) using an objective lens of 100 $\times$  with a 532 nm excitation source. In UV light irradiation, the UT and UTG core-shell samples on the Si/SiO<sub>2</sub> substrates were placed below a UV lamp source with a distance of 5 cm. Procedures of preparation and RhB degradation of 3D urchin-like TiO<sub>2</sub>@graphene SERS substrates under irradiation of UV light are displayed in Scheme 1. The adsorption and photocatalytic activities of the samples were also measured *via* UV-visible absorption spectrophotometer (UV-2450, UV/Vis spectrophotometer, Shimadzu, Japan).

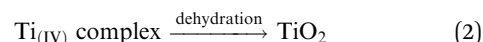
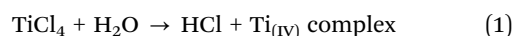
## 3. Results and discussion

### 3.1. The morphology, structure and composition characterization of 3D urchin-like TiO<sub>2</sub>@Gr core-shell

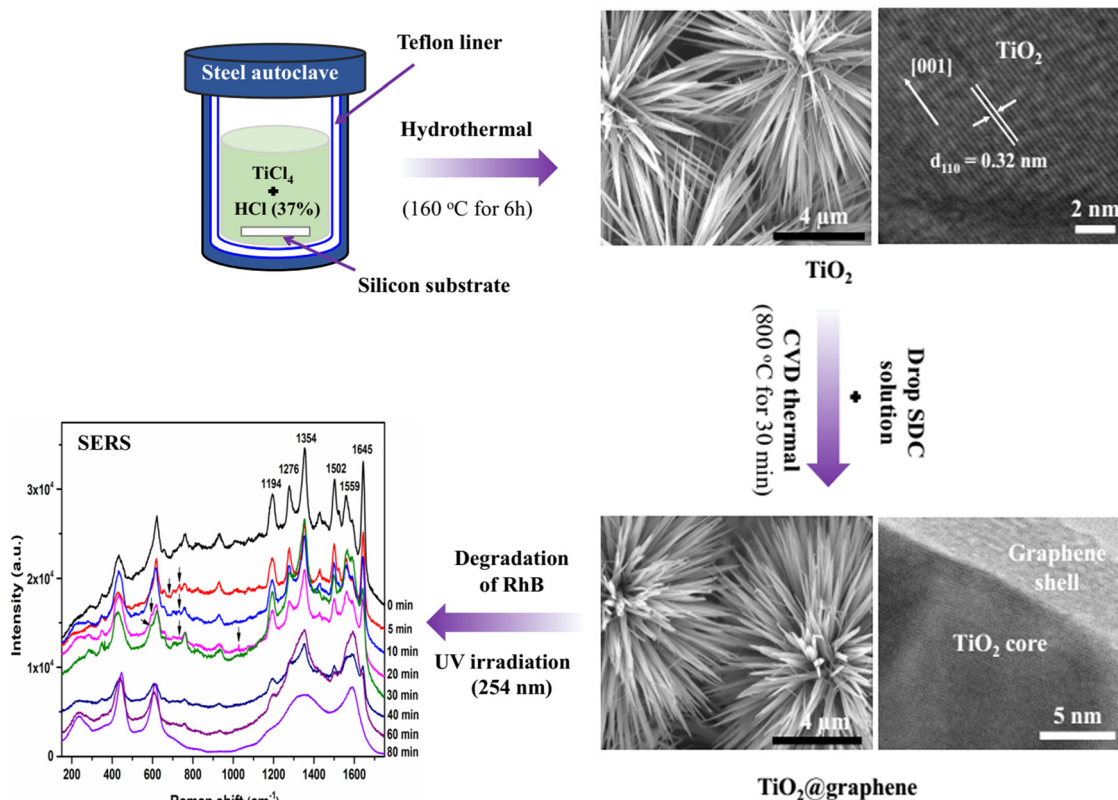
The morphology, size and microstructure of 3D urchin-like TiO<sub>2</sub> (UT) and 3D urchin-like TiO<sub>2</sub>@graphene (UTG) core-shell

were observed by SEM and TEM images, as shown in Fig. 1a–d. It is observed that UTs are covered by numerous nanoneedles, which grow radially from the core of the microspheres and grow uniformly on the surface of the Si/SiO<sub>2</sub> substrate. The diameter range of the UTs is from 10  $\mu\text{m}$  to 12  $\mu\text{m}$ , as shown in Fig. 1a. The thickness of the UT layer covered on the surface of the Si/SiO<sub>2</sub> substrate is about 65  $\mu\text{m}$  (Fig. 1b). Fig. 1c and d are high-magnification SEM images of UTs and UTGs, respectively. Fig. 1c and d indicate that the surface of the UT and UTG microspheres is very smooth and clean. The HRTEM image in the inset of Fig. 1c shows that the distance between the adjacent lattice fringes of the UT crystals is about 0.32 nm. This distance is attributed to the interplanar distance of rutile TiO<sub>2</sub> (110).<sup>18,28–30</sup> The nanoneedles are grown along the rutile (110) crystal plane, and with a preferred orientation in the [001] direction.<sup>29,30</sup>

The hydrothermal reactions of the growth of the rutile TiO<sub>2</sub> nanoneedles were as follows:<sup>31,32</sup>

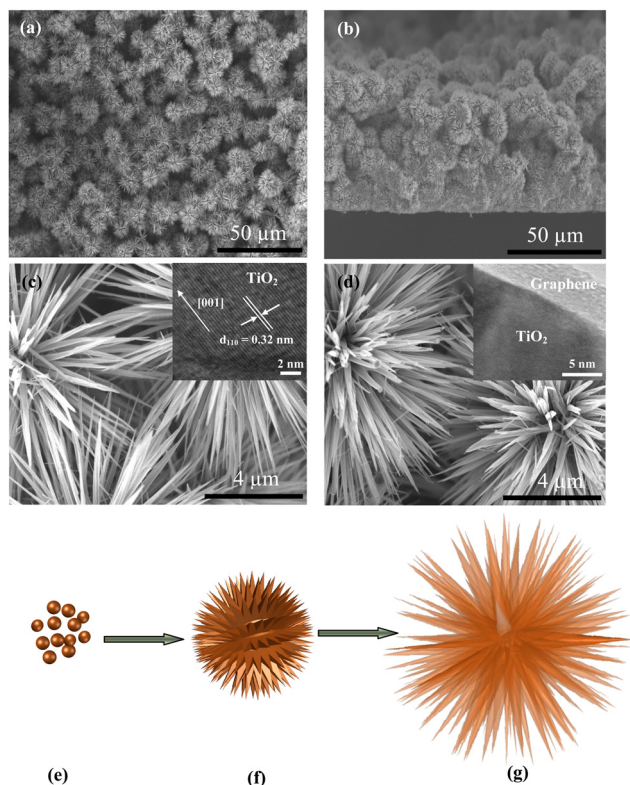


The presence of abundant amounts of Ti<sup>4+</sup> precursor, H<sup>+</sup>, and Cl<sup>−</sup> ions contributes to the formation of rutile TiO<sub>2</sub>. The presence of highly acidic conditions may reduce the hydrolysis of the Ti source, and the selective adsorption of Cl<sup>−</sup> on rutile



Scheme 1 Fabrication procedure and RhB degradation of TiO<sub>2</sub>@graphene SERS substrates under the irradiation of UV light.





**Fig. 1** Surface and cross sectional SEM images (a) and (b) of UTs on the SiO<sub>2</sub> substrate, high-magnification SEM images and the insets are HRTEM images of (c) UTs and (d) UTGs. Schematic illustration of the formation process of the UTs: (e) nucleation, (f) dissolution and recrystallization, and (g) self-assembly.

(110) can promote the anisotropic growth along the [001] direction of rutile TiO<sub>2</sub>.<sup>31,32</sup>

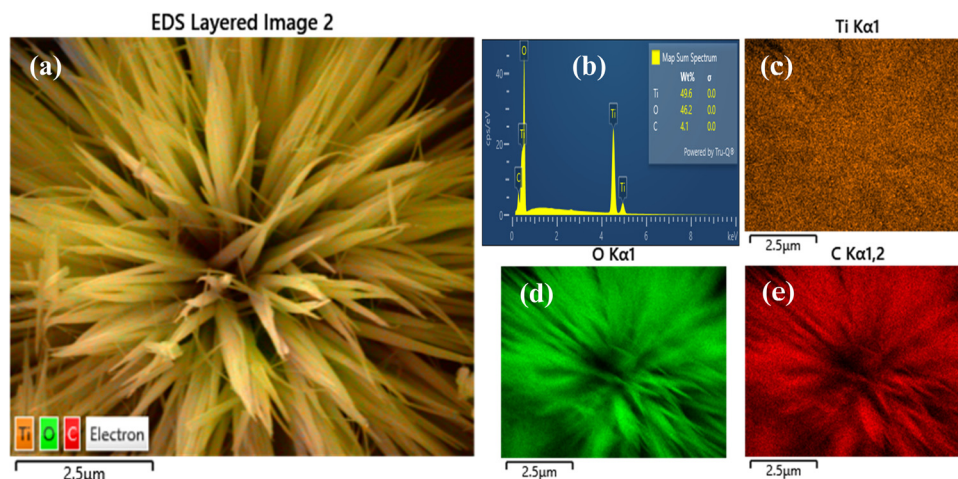
As reported previously,<sup>29,30</sup> the mechanism that is responsible for the formation of the 3D urchin-like TiO<sub>2</sub> (UT) structure may follow a process of nucleation, dissolution and recrystallization, and self-assembly, as shown in Fig. 1e–g. At the first stage, amorphous or rutile nuclei of TiO<sub>2</sub> were formed, and

then they coalesced into crinkly spherical units as shown in Fig. 1e. These TiO<sub>2</sub> spherical units dissolved and recrystallized to gradually form TiO<sub>2</sub> nanoneedles on the surface of the microspheres (Fig. 1f). Finally, the 3D urchin-like TiO<sub>2</sub> structures were formed due to the directional self-assembly of the TiO<sub>2</sub> nanoneedles (Fig. 1g).

The thickness of graphitic layer covered TiO<sub>2</sub> was provided from the HRTEM image. As shown in the HRTEM image in the inset of Fig. 1d, a very thin layer graphitic structure was observed on the surface of the UT. The thickness of the graphitic layer was about several nm. Furthermore, the presence of graphitic layer covered TiO<sub>2</sub> was confirmed by the energy dispersive X-ray spectroscopy (EDS) elemental mapping (Fig. 2). The EDS results (Fig. 2a) show that the UTG was composed of Ti, O and C elements without other elements, and the corresponding contents are 49.6 wt%, 46.2 wt% and 4.1 wt%, respectively. Fig. 2b and c suggests that the Ti (orange), O (green), and C (red) elements were uniformly distributed in the UTGs.

Fig. 3a and b show XRD patterns and Raman spectra of the UT and UTG samples, respectively. As shown in Fig. 3a, the prepared UT exhibits main diffraction peaks at  $2\theta = 27.5^\circ$ ,  $36.1^\circ$ ,  $39.2^\circ$ ,  $41.3^\circ$ ,  $44.1^\circ$ ,  $54.4^\circ$ ,  $56.7^\circ$ ,  $62.8^\circ$ , and  $64.1^\circ$  corresponding to the (1 1 0), (1 0 1), (200), (111), (2 1 0), (2 1 1), (2 2 0), (002), and (310) planes of TiO<sub>2</sub> rutile (JCPDS no. 01-087-0710).<sup>28</sup> The peaks are sharp and intense, indicating good crystallinity of TiO<sub>2</sub>. For UTG, the XRD pattern also shows the main diffraction peaks of TiO<sub>2</sub> crystals with rutile phase. However, the diffraction peaks of graphitic carbon cannot be detected in the XRD patterns (Fig. 3a). This is possibly due to the small amount of graphitic carbon grown onto the surface of UT and/or due to the strong diffraction of rutile TiO<sub>2</sub>.

Fig. 3b presents the Raman spectra of the UT and UTG samples. As shown in Fig. 3b, the prepared UT shows the presence of three peaks at  $607\text{ cm}^{-1}$ ,  $442\text{ cm}^{-1}$ , and  $235\text{ cm}^{-1}$  corresponding to the rutile phase.<sup>33</sup> The Raman spectrum of the UTG sample after growth of graphitic carbon also shows three rutile peaks of TiO<sub>2</sub> and three important graphitic bands at  $\sim 1330\text{ cm}^{-1}$ ,  $\sim 1600\text{ cm}^{-1}$ , and  $\sim 2630\text{--}2640\text{ cm}^{-1}$



**Fig. 2** (a) SEM image, (b) EDS spectrum, and (c)–(e) elemental mapping images of (c) Ti, (d) O and (e) C elements of UTG.



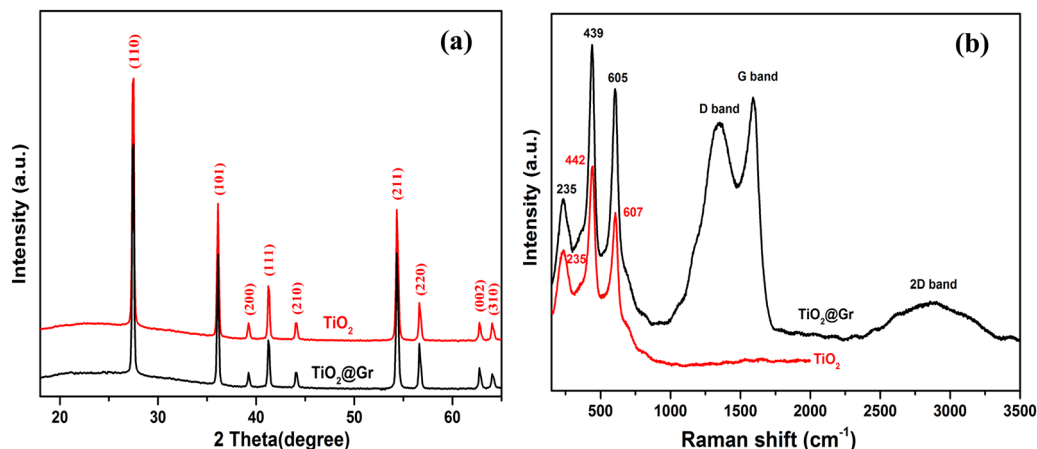


Fig. 3 (a) XRD and (b) Raman spectra of UT and UTG samples.

corresponding to the D band, G band, and 2D band, respectively. The D band and G band are characteristics of structure defects and  $sp^2$ -hybridization of carbon atoms, respectively. The 2D band, the second-order Raman of the D band, is the characteristic band corresponding to the number of graphitic sheets that form the multi-layer graphene structure.<sup>18,34–38</sup> These important characteristic peaks of graphitic confirmed the presence of graphene (Gr) structures in the crystalline structure of  $TiO_2$  rutile. Moreover, compared to the Raman spectrum of the UT sample, there is a red-shift of about  $2\text{ cm}^{-1}$  for the  $605\text{ cm}^{-1}$  band and about  $3\text{ cm}^{-1}$  for the  $439\text{ cm}^{-1}$  band. These shifts can be assigned to interactions of  $TiO_2$  with Gr during the thermal CVD process.<sup>18</sup>

### 3.2. Optical characterizations of 3D urchin-like $TiO_2@Gr$

The optical characteristics of UT and UTG samples were measured by UV-vis absorption and photoluminescence (PL) spectra. Fig. S1a and b (ESI<sup>†</sup>) illustrate the absorption spectra and optical band gap of the UT and UTG samples, respectively. As shown in Fig. S1a (ESI<sup>†</sup>), the absorption edges of the UTG display a red-shift compared with UT. These shifts can be assigned to restoration of the  $\pi$ - $\pi$  conjugation network of the Gr in the hybrid.<sup>39</sup> Moreover, the incorporation of Gr into the  $TiO_2$  system also reduced the energy band gap of the photocatalyst. According to the Tauc equation, the energy band gaps of UT and UTG were calculated to be around 2.84 eV and 2.77 eV, respectively. In order to study the recombination of photo-induced electron-hole pairs, photoluminescence (PL) measurements were conducted. As shown in Fig. S1c (ESI<sup>†</sup>), the PL intensity of UTG was lower than that of UT. This attenuation is attributed to Gr's exceptional electron transport properties, which effectively suppress recombination of electron-hole pairs, hence enhancing the photocatalytic degradation efficiency of the organic contaminants.<sup>39</sup>

### 3.3. SERS characterizations of 3D urchin-like $TiO_2@Gr$

For comparison, the SERS performance of the UT and UTG samples were studied under the same conditions. The Raman spectra of the RhB molecule with a concentration of  $10^{-5}\text{ M}$  on the UT and UTG samples are shown in Fig. 4. For the SERS

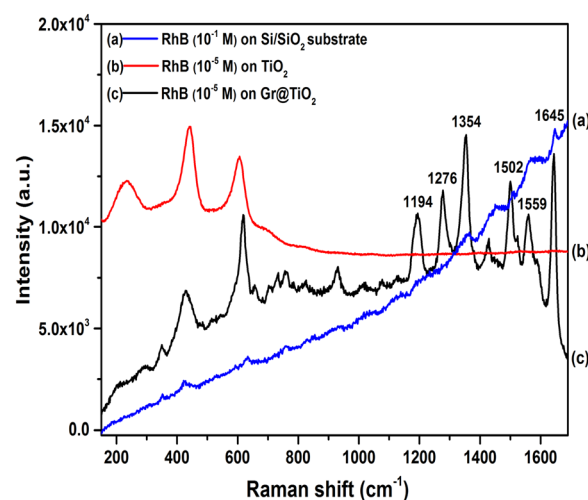


Fig. 4 SERS spectra of RhB adsorbed on: (a) a  $Si/SiO_2$  sample with the concentration of  $10^{-1}\text{ M}$ , and (b) UT, and (c) UTG samples with an RhB concentration of  $10^{-5}\text{ M}$ .

substrate of the UT sample, no Raman scattering signals of RhB molecules are detected. For the SERS substrate of the UTG sample, the SERS spectrum shows strong signals of RhB molecules at  $1645\text{ cm}^{-1}$ ,  $1559\text{ cm}^{-1}$ ,  $1502\text{ cm}^{-1}$ ,  $1354\text{ cm}^{-1}$ ,  $1276\text{ cm}^{-1}$ , and  $1194\text{ cm}^{-1}$ . The RhB bands at  $1645\text{ cm}^{-1}$ ,  $1559\text{ cm}^{-1}$ ,  $1502\text{ cm}^{-1}$ , and  $1354\text{ cm}^{-1}$  are assigned to aromatic C-C stretching modes;<sup>40</sup> the band at  $1276\text{ cm}^{-1}$  is assigned to C-H bending vibration, and the band at  $1194\text{ cm}^{-1}$  is attributed to aromatic C-C stretching.<sup>41–43</sup> The SERS properties show the remarkably improved SERS activity of the RhB molecules over the UTG compared to UT. The SERS sensitivity of the UTG substrate can be evaluated by calculating the enhancement factor (EF), using the following equation:<sup>6,44–46</sup>

$$EF = (I_{\text{SERS}}/I_{\text{Ref}}) \times (C_{\text{Ref}}/C_{\text{SERS}}) \quad (3)$$

where  $I_{\text{SERS}}$  and  $I_{\text{Ref}}$  are the Raman intensities at  $1645\text{ cm}^{-1}$  of RhB molecules adsorbed on the  $TiO_2@Gr$ /SERS substrates-coated graphene (UTG) and non-SERS substrate ( $SiO_2$ ).  $C_{\text{SERS}}$  and  $C_{\text{Ref}}$  are



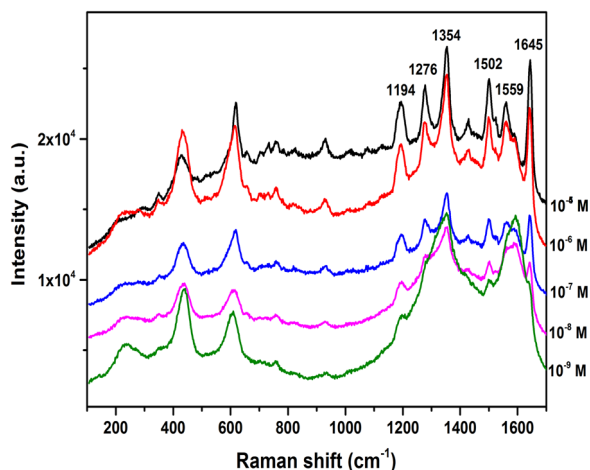


Fig. 5 SERS spectra of RhB adsorbed on UTG samples with the concentration from  $10^{-5}$  M to  $10^{-9}$  M of RhB.

the corresponding concentration of the RhB molecule dropped on the SERS substrate ( $10^{-5}$  M) and non-SERS substrate ( $10^{-1}$  M). The EF of the UTG SERS substrate calculated for the strongest peak at  $1645\text{ cm}^{-1}$  was  $1.3 \times 10^5$  of RhB molecules with a concentration of  $10^{-5}$  M.

Fig. 5 shows the SERS spectra of RhB molecules with concentrations ranging from  $10^{-5}$  M to  $10^{-9}$  M. It can be found that the Raman signal intensities of RhB molecules decrease with decreasing probe molecule concentration and the minimum detection concentration of  $10^{-9}$  M can be detected. The possible charge transfer mechanism for the enhancement effect in the SERS spectra of UTG for the RhB molecules is demonstrated in Fig. 6. The presence of graphene (Gr) in the UTG hybrid structure acts as the electron acceptor and transporter. The excited electrons of  $\text{TiO}_2$  can be transferred into Gr from the surface state energy levels ( $E_{ss}$ ) of  $\text{TiO}_2$ . The Gr acts as the electron acceptor of  $\text{TiO}_2$  and prevents the charge recombination, and provides an indirect charge transfer from  $\text{TiO}_2$  to RhB molecules. With excellent conductivity, Gr can promote the transport of charge carriers and the charge transfer efficiency from  $\text{TiO}_2$  to RhB molecules. A strong interfacial charge transfer between  $\text{TiO}_2$  and Gr would enhance the SERS signals of the RhB molecules.

### 3.4. Photocatalytic activity of a 3D urchin-like $\text{TiO}_2$ @Gr core-shell

The photo-catalytic activity of UTGs was studied by detecting the SERS spectra of RhB molecules. The SERS spectra of the RhB molecules with concentration of  $10^{-5}$  M after the decomposition by UTGs under different UV irradiation times of 0–80 min are shown in Fig. 7a. It can be found that the characteristic Raman peaks of the RhB molecules were gradually decreased with increase of the UV irradiation time. During the degradation process, besides the appearance of the main characteristic Raman peaks of RhB molecules, some strange peaks were also observed at different times, such as  $593.8\text{ cm}^{-1}$  (20 min, and 30 min),  $732.6\text{ cm}^{-1}$  (5 min, 10 min, and 20 min),

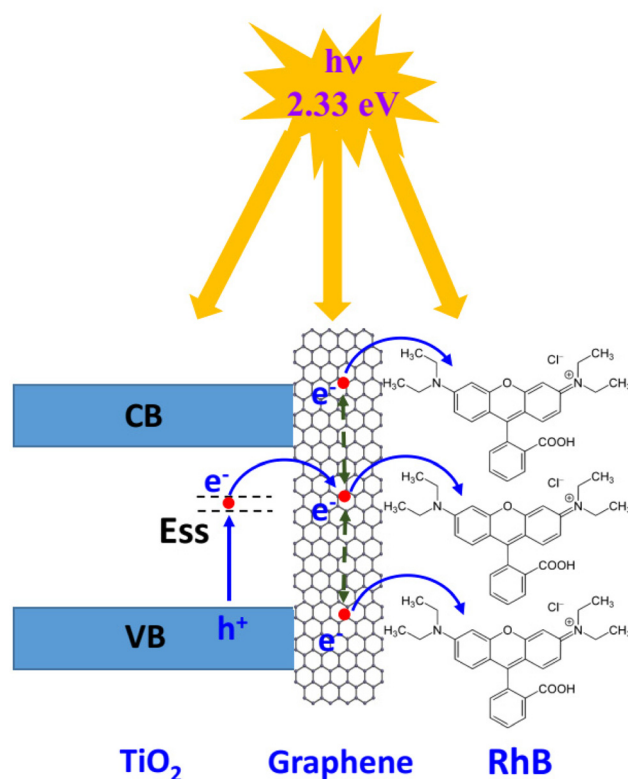


Fig. 6 Possible charge transfer mechanism of  $\text{TiO}_2$ @graphene/RhB structure.

and  $1028.5\text{ cm}^{-1}$  (20 min). These strange peaks could be contributed from the intermediates formed at the interface of the UTGs and air. After the UV irradiation time of 80 min, almost all the peaks of RhB molecules and intermediates had disappeared completely.

The photo-catalytic efficiency of UTs and UTGs was studied through the RhB Raman peak intensity located at  $1645\text{ cm}^{-1}$ . This is a peak which has the strongest intensity of RhB in the SERS spectra. The photocatalytic rate ( $k$ ) was evaluated *via*  $\ln(P/P_0)$  versus the irradiation time (Fig. 7b), where  $P_0$  and  $P$  are the intensity of the Raman peaks of RhB initially and after different UV irradiation times, respectively. From the slope of the fitting curve (Fig. 7b), the photocatalytic rate of UTGs is about  $0.069\text{ min}^{-1}$ . The photo-catalytic activity of the UTG SERS substrate for RhB degradation under UV irradiation was also compared with previously reported SERS substrates for organic pollutant degradation under different irradiation sources (Table 1).

The photo-catalytic activity of the UTs and UTGs was also studied by UV-visible absorption spectrophotometry (Fig. S2, ESI†). Compared to UV-visible absorption spectrophotometry, the SERS technique can detect the  $k$  values of the chemical bonds in the RhB structural changes during the course of RhB degradation at the molecular level, while UV-visible absorption spectrophotometry can detect only the  $k$  value of the RhB in the bulk solution. To identify the radical scavengers generated during the photocatalysis reaction, three different capturing agents were introduced into the reaction mixture, namely





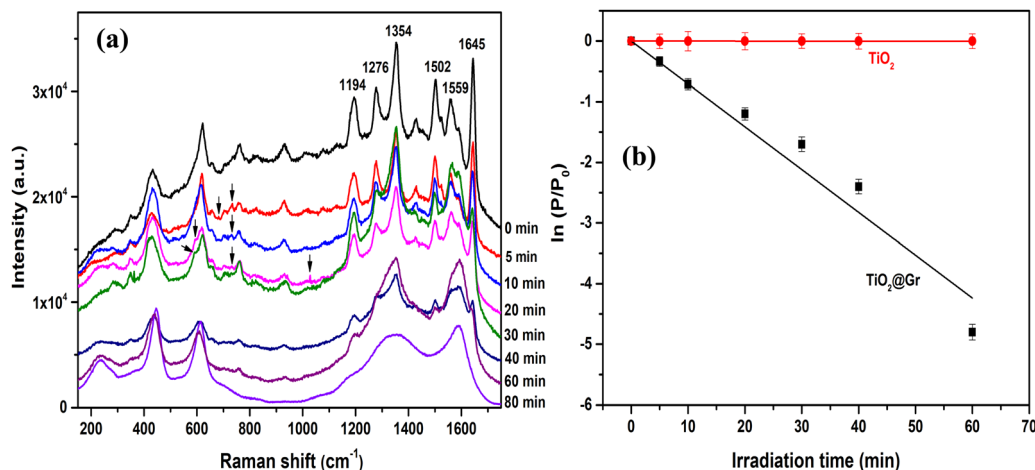


Fig. 7 (a) SERS spectra and (b) linear relation between logarithm of SERS intensities ( $\ln(P/P_0)$ ) at  $1645\text{ cm}^{-1}$  versus irradiation time (min) for the degradation of RhB.

Table 1 Comparison of photocatalytic activity from different SERS substrates for organic pollutant degradation

SERS substrate for photocatalytic degradation	Probe molecule	Irradiation source	Complete degradation time (min)	Concentration ( $\text{mol L}^{-1}$ )	Degradation rate ( $k$ ) $\text{min}^{-1}$	Ref.
Ag-G-TiO <sub>2</sub> /G-Ag-TiO <sub>2</sub> /G-Ag-TiO <sub>2</sub> -G	R6G	UV ( $\lambda > 410\text{ nm}$ )	60	$10^{-6}$	0.0371/ 0.0301/ 0.0111	11
Ag/TiO <sub>2</sub> NTs	R6G	UV (254 nm)	140	$10^{-6}$	0.02479	47
TiO <sub>2</sub> /AgNP	R6G	UV (385 nm)	120	$10^{-6}$	0.029	48
Ag-TiO <sub>2</sub> nanorods	R6G	UV (254 nm)	90	$10^{-5}$	0.043	49
g-C <sub>3</sub> N <sub>4</sub> /AuNPs	R6G	Xenon lamp	20	$10^{-5}$	0.1102	50
Ag-TiO <sub>2</sub> nanotube arrays	R6G	UV (365 nm)	150	$10^{-7}$	NA	51
ZnO/Ag	RhB	Xenon lamp, 420 nm	150	$10^{-5}$	NA	43
Ag/PDA/ZnO@GMF	RhB	UV (254 nm)	40	$10^{-3}$	0.0391	52
Ag/TiO <sub>2</sub>	RhB	LED light (370 nm)	120	$10^{-3}$	NA	53
TiO <sub>2</sub> /Ag	RhB	LED light (370 nm)	90	$10^{-3}$	NA	54
Ag <sub>2</sub> CO <sub>3</sub>	RhB	Mercury-xenon lamp light	30	$10^{-5}$	NA	55
CNs/AgNSSs@CNQDs-GLs membrane	RhB	Xenon lamp (400 nm)	80	$10^{-7}$	NA	56
Ag/Cu <sub>2</sub> O	RhB	Xenon lamp	120	$10^{-7}$	NA	57
Ag@SiO <sub>2</sub> /CNNDs/EGLs	RhB	Xenon lamp (400 nm)	20	$10^{-6}$	NA	58
TiO <sub>2</sub> @Gr core-shell	RhB	UV (254 nm)	80	$10^{-5}$	0.069	This work

PDA, polydopamine; glass microfibre filters, GMF; CNs, graphitic-carbon-nitride nanosheets; CNQDs, graphitic-carbon-nitride quantum-dots (CNQDs-GLs); AgNSSs, Ag nanospheres; CNNDs, g-C<sub>3</sub>N<sub>4</sub> nanodots; EGLs, exfoliated graphene layers.

ethylenediaminetetraacetic acid disodium (EDTA),<sup>59–61</sup> isopropyl alcohol (IPA),<sup>59–63</sup> and benzoquinone (BQ).<sup>59–62</sup> EDTA, IPA, and BQ scavengers were used to trap holes ( $h^+$ ), hydroxyl radicals ( $\bullet\text{OH}$ ), and superoxide radicals ( $\bullet\text{O}_2^-$ ), respectively. As depicted in Fig. S3 (ESI<sup>†</sup>), the photocatalytic degradation of RhB by UTGs was found to decrease from 98.5% to 40.2%, 57.3%, and 64.6% in the presence of EDTA, BQ, and IPA, respectively. The  $k$  value for the photocatalytic degradation of RhB by UTGs was found to decrease from  $0.046\text{ min}^{-1}$  to ( $0.006$ – $0.014\text{ min}^{-1}$ ) in the presence of scavengers under UV light irradiation. The order of inhibition of the photodegradation efficiency was EDTA ( $h^+$ ) > BQ ( $\bullet\text{O}_2^-$ ) > IPA ( $\bullet\text{OH}$ ). These results suggest that  $h^+$  and  $\bullet\text{O}_2^-$  radicals play the main role for the decomposition of RhB, and  $\bullet\text{OH}$  plays a relatively minor role in RhB photodecomposition.

The photocatalytic degradation mechanism for UTGs towards RhB is shown in Fig. 8. Under UV light irradiation, the excited electrons from the surface of TiO<sub>2</sub> are transferred from the valence band (VB) of TiO<sub>2</sub> to the conduction band (CB), to thus generate electron-hole ( $e^-$ - $h^+$ ) pairs (eqn (4)). The photogenerated  $e^-$  in the CB of TiO<sub>2</sub> can be transferred to Gr through the intimate interfacial contact between TiO<sub>2</sub> and Gr (eqn(5)). With good conductivity, Gr will enhance the separation and lifetime of the  $e^-$ - $h^+$  pairs and thus will improve the photocatalytic activity of the UTGs. The electrons on the surface of Gr react with dissolved oxygen ( $\text{O}_2$ ) to form radical anions  $\bullet\text{O}_2^-$  (eqn (6)). The photogenerated  $e^-$  on the TiO<sub>2</sub> surface could also be trapped directly by the  $\text{O}_2$  to form radical anions  $\bullet\text{O}_2^-$ . These radical anions  $\bullet\text{O}_2^-$  combine with  $\text{H}_2\text{O}$  molecules to form  $\bullet\text{OH}$  radicals (eqn (7)–(9)). Meanwhile, the separated holes are

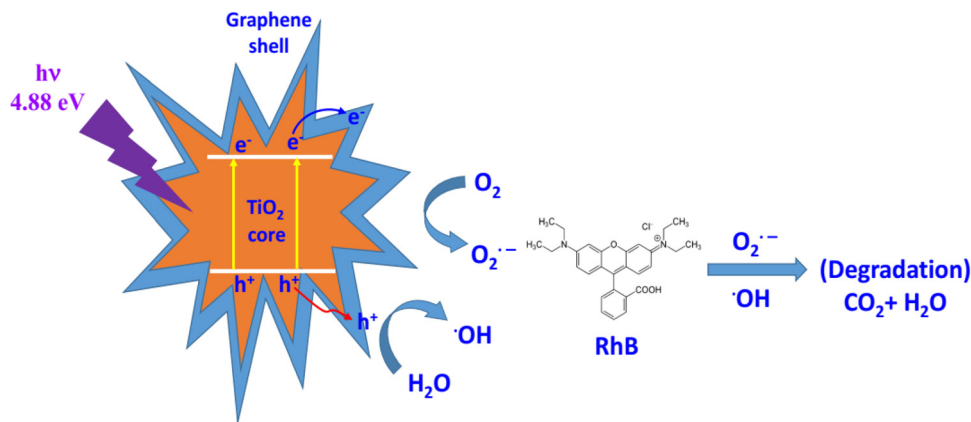
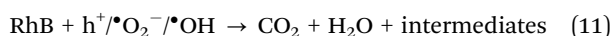
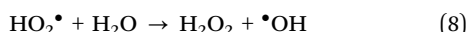
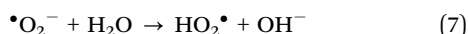
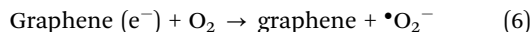
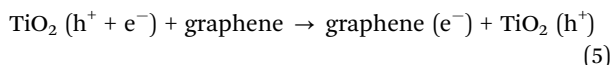
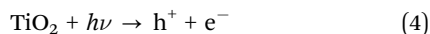


Fig. 8 Degradation mechanism for UTG towards RhB.

consumed to oxidize the dye molecules directly. On the other hand, holes from the VB of TiO<sub>2</sub> can react with either H<sub>2</sub>O or OH<sup>-</sup> ions adsorbed on the catalyst surface to generate highly reactive •OH radicals (eqn (10)). The holes, •O<sub>2</sub><sup>-</sup> and •OH activated radicals degrade RhB molecules into CO<sub>2</sub>, H<sub>2</sub>O and other intermediates (eqn (11)).<sup>64</sup> The major reaction equations that are involved in the photocatalytic degradation mechanism for UTG towards RhB are described by eqn (4)–(11).



## 4. Conclusions

In this work, 3D urchin-like core-shell TiO<sub>2</sub>@Gr (UTG) was successfully fabricated on Si/SiO<sub>2</sub> substrates by a simple hydrothermal combined with thermal CVD method using sodium deoxycholate (SDC) surfactant as a carbon source. The fabricated UTG samples were used as SERS substrates to detect and decompose rhodamine B (RhB) under UV irradiation. It was found that the Gr shell played an important role in the properties of the UTG substrate. The minimum detectable concentration of RhB on the UTG SERS substrate can reach  $1 \times 10^{-9}$  M with an enhancement factor of  $1.3 \times 10^5$ . The graphene acts as the electron acceptor of TiO<sub>2</sub> and prevents charge recombination, and provides indirect charge transfer from TiO<sub>2</sub> to RhB molecules. Moreover, the UTG SERS substrate can almost completely decompose RhB with a degradation rate of  $0.069 \text{ min}^{-1}$

under UV 254 nm irradiation within 80 min. The photocatalytic degradation mechanism for the UTGs towards RhB was also presented in detail. The results show that the UTG SERS substrates can be further employed for detecting substances and degrading hazardous pollutants.

## Author contributions

Nguyen Thi Huyen: methodology, investigation, formal analysis, data curation, conceptualization. Tran Ai Suong Suong, Cao Thi Thanh, Pham Van Trinh, Nguyen Van Tu, Bui Hung Thang, Tran Van Hau: methodology, investigation, formal analysis. Pham Thanh Binh, Vu Duc Chinh, Pham Van Hai, Vu Xuan Hoa: investigation, formal analysis. Tran Van Tan: investigation, formal analysis. Phan Ngoc Minh: supervision, investigation. Hiroya Abe: supervision, investigation. Nguyen Van Chuc: writing – review & editing, writing – original draft, validation, supervision, project administration, methodology, investigation, formal analysis, conceptualization. All authors discussed and approved the paper.

## Data availability

The datasets used and/or analyzed during the current study are available from the corresponding author upon reasonable request. In addition, all the data generated or analyzed during this study are included in this article.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This work was financially supported by the Vietnam Academy of Science and Technology (VAST) under Project NCXS02.03/24-25.





## References

- 1 J. Tong, H. Zhai, S. Zhao, L. Song, G. Wang, N. Feng, P. Tan, J. Xie and J. Pan, *J. Colloid Interface Sci.*, 2024, **653**, 285–295.
- 2 M. Rani, Keshu and U. Shanker, *Int. J. Environ. Sci. Technol.*, 2023, **20**, 5491–5508.
- 3 X. Song, J. He, M. Zhang, S. Zhang and J. Liu, *Colloids Surf. A Physicochem. Eng. Asp.*, 2024, **684**, 133116.
- 4 P. Butmee, A. Samphao and G. Tumcharern, *J. Hazard. Mater.*, 2022, **437**, 129344.
- 5 L. Tan, Y. Cao, J. Yan, K. Mao, L. Liu, X. Wang, W. Ye, R. A. Harris and H. Zhang, *Anal. Chim. Acta*, 2024, **1287**, 342047.
- 6 R. Ziad, S. Columbus, A. Elgamouz, K. Daoudi, A. N. Kawde, K. Ramachandran and M. Gaidi, *Spectrochim. Acta Part A Mol. Biomol. Spectrosc.*, 2023, **297**, 122701.
- 7 Y. Wang, M. Zhang, H. Yu, Y. Zuo, J. Gao, G. He and Z. Sun, *Appl. Catal., B*, 2019, **252**, 174–186.
- 8 Z. Li, K. Han, A. Zhang, T. Wang, Z. Yan, Z. Ding, Y. Shen, M. Zhang and W. Zhang, *Talanta*, 2024, **266**, 125070.
- 9 B. Zhao, H. Liu, L. Xia, Z. Wang and C. Zhang, *ACS Appl. Nano Mater.*, 2022, **5**, 15341–15352.
- 10 X. Jiang, D. Yin, M. Yang, J. Du, W. Wang, L. Zhang, L. Yang, X. Han and B. Zhao, *App. Surf. Sci.*, 2019, **487**, 938–944.
- 11 X. Zhang, N. Wang, R. Liu, X. Wang, Y. Zhu and J. Zhang, *Opt. Mater. Express*, 2018, **8**, 704–717.
- 12 M. Zhang, H. Sun, X. Chen, H. Zhou, L. Xiong, W. Chen, Z. Chen, Z. Bao and Y. Wu, *J. Alloy. Compd.*, 2021, **864**, 158189.
- 13 K. P. Cabral, W. Kurniawan and H. Hinode, *J. Environ. Chem. Eng.*, 2015, **3**, 2786–2796.
- 14 H. Li, B. Yang, B. Yu, N. Huang, L. Liu, J. Lu and X. Jiang, *App. Surf. Sci.*, 2021, **541**, 148486.
- 15 W. Liu, Z. Wang, W. Yan, Z. Zhao, L. Shi, L. Huang, Y. Liu, X. He and S. Cui, *Carbon*, 2023, **202**, 389–397.
- 16 X. Ling, L. G. Moura, M. A. Pimenta and J. Zhang, *J. Phys. Chem. C*, 2012, **116**, 25112–25118.
- 17 L. L. Qu, N. Wang, Y. Y. Li, D. D. Bao, G. H. Yang and H. T. Li, *J. Mater. Sci.*, 2017, **52**, 8311–8320.
- 18 A. R. Biris, D. Toloman, A. Popa, M. D. Lazar, G. K. Kannarpady, V. Saini, F. Watanabe, B. P. Chhetri, A. Ghosh and A. S. Biris, *Phys. E*, 2016, **81**, 326–333.
- 19 M. A. Fitri, M. Ota, Y. Hirota, Y. Uchida, K. Hara, D. Ino and N. Nishiyama, *Mater. Chem. Phys.*, 2017, **198**, 42–48.
- 20 H. Liu, Z. Chen, L. Zhang, D. Zhu, Q. Zhang, Y. Luo and X. Shao, *J. Phys. Chem. C*, 2018, **122**, 6388–6396.
- 21 S. Karamat, U. Khalique, A. Usman, A. Javaid and A. Oral, *Arab J. Sci. Eng.*, 2022, **47**, 7779–7788.
- 22 G. Lui, J. Y. Liao, A. Duan, Z. Zhang, M. Fowler and A. Yu, *J. Mater. Chem. A*, 2013, **1**, 12255.
- 23 X. Jiang and J. Wang, *App. Surf. Sci.*, 2023, **612**, 155849.
- 24 Z. Zhou, R. Yang, Y. Teng, Y. Li and M. Wei, *J. Colloid Interface Sci.*, 2023, **637**, 533–540.
- 25 N. G. Nair, V. G. Gandhi, K. Modi and A. Shukla, *Mater. Today Proc.*, 2024, DOI: [10.1016/j.matpr.2023.12.049](https://doi.org/10.1016/j.matpr.2023.12.049).
- 26 D. Heltina, D. I. Mastura, A. Amri and M. P. Sembiring, *Mater. Today Proc.*, 2023, **87**, 293–298.
- 27 A. B. Migdadi, Q. M. Al-Bataineh, A. A. Ahmad, H. M. Al-Khateeb and A. Telfah, *J. Alloy. Compd.*, 2024, **971**, 172794.
- 28 S. Wannapop and A. Somdee, *Ceram. Int.*, 2020, **46**, 25758–25765.
- 29 L. Xiang, X. Zhao, J. Yin and B. Fan, *J. Mater. Sci.*, 2012, **47**, 1436–1445.
- 30 C. Shi, M. Egi, J. Shi, Z. Huang and H. Qi, *J. Colloid Interface Sci.*, 2023, **650**, 1736–1748.
- 31 A. Kumar, A. R. Madaria and C. Zhou, *J. Phys. Chem. C*, 2010, **114**, 7789.
- 32 M. Lv, D. Zheng, M. Ye, L. Sun, J. Xiao, W. Guo and C. Lin, *Nanoscale*, 2012, **4**, 5872–5879.
- 33 Y. Jiang, W. Zhao, S. Li, S. Wang, Y. Fan, F. Wang, X. Qiu, Y. Zhu, Y. Zhang, C. Long and Z. Tang, *J. Am. Chem. Soc.*, 2022, **144**, 15977–15987.
- 34 C. T. Thanh, N. T. Huyen, P. V. Trinh, N. V. Tu, V. T. Thu, V. C. Tu, D. N. Nhiem, P. T. Binh, N. N. Anh, V. X. Hoa, P. N. Minh, H. Abe and N. V. Chuc, *Diamond Relat. Mater.*, 2025, **152**, 111889.
- 35 P. N. D. Duoc, N. H. Binh, T. V. Hau, C. T. Thanh, P. V. Trinh, N. V. Tuyen, N. V. Quynh, N. V. Tu, V. D. Chinh, V. T. Thu, P. D. Thang, P. N. Minh and N. V. Chuc, *J. Hazard. Mater.*, 2020, **400**, 123185.
- 36 C. T. Thanh, P. N. D. Duoc, P. V. Trinh, N. T. Huyen, N. V. Tu, C. T. Anh, P. V. Hai, K. Yoshida, H. Abe and N. V. Chuc, *Mater. Lett.*, 2023, **330**, 133308.
- 37 S. Kumar, R. D. Kaushik, G. K. Upadhyay and L. P. Purohit, *J. Hazard. Mater.*, 2021, **406**, 124300.
- 38 C. T. Thanh, N. T. Huyen, V. T. Thu, P. V. Trinh, N. V. Tu, B. H. Thang, T. V. Hau, D. Tuan, M. T. Phuong, P. T. Binh, P. N. Minh, H. Abe and N. V. Chuc, *Mater. Lett.*, 2025, **386**, 138209.
- 39 Y. Wang, Z. Mo, P. Zhang, C. Zhang, L. Han, R. Guo, H. Gou, X. Wei and R. Hu, *Mater. Des.*, 2016, **99**, 378–388.
- 40 J. Wang, L. Jiang, F. Liu, M. Jia, M. Liu, J. Li and Y. Lai, *Chem. Eng. J.*, 2021, **407**, 127195.
- 41 L. T. Q. Ngan, D. T. Cao, C. T. Anh, K. N. Minh and L. V. Vu, *J. Electron. Mater.*, 2019, **48**, 5328–5332.
- 42 P. Dai, Y. Xue, X. Wang, Q. Weng, C. Zhang, X. Jiang, D. Tang, X. Wang, N. Kawamoto, Y. Ide, M. Mitome, D. Golberg and Y. Bando, *Nanoscale*, 2015, **7**, 18992–18997.
- 43 X. Chen, L. Zhu, Z. Ma, M. Wang, R. Zhao, Y. Zou and Y. Fan, *Nanomaterials*, 2022, **12**, 2394.
- 44 E. C. Le Ru, E. Blackie, M. Meyer and P. G. Etchegoin, *J. Phys. Chem. C*, 2007, **111**, 13794–13803.
- 45 X. Y. Wang, W. Y. Zhang, Y. J. Hu, H. Y. Song, A. Zeeshan, C. Ge and S. B. Liu, *Spectrochim. Acta, Part A*, 2024, **310**, 123932.
- 46 T. B. Pham, T. H. C. Hoang, V. H. Pham, V. C. Nguyen, T. V. Nguyen, D. C. Vu, V. H. Pham and H. Bui, *Sci. Rep.*, 2019, **9**, 12590.
- 47 X. Chong, B. Zhao, R. Li, W. Ruan and X. Yang, *Colloids Surf. A Physicochem. Eng. Asp.*, 2015, **481**, 7–19.
- 48 H. Y. Wu, H. C. Lin, Y. H. Liu, K. L. Chen, Y. H. Wang, Y. S. Sun and J. C. Hsu, *Molecules*, 2022, **27**, 6755.
- 49 S. Kumar, D. K. Lodhi and J. P. Singh, *RSC Adv.*, 2016, **6**, 45120.



- 50 W. Chen, W. Wang, H. Xing, W. Li, H. He, W. Li, P. K. Chu, G. Song, H. Wang and P. Li, *Surf. Interfaces*, 2023, **38**, 102808.
- 51 H. Zhai, C. Zhu, X. Wang, Y. Yan and H. Tang, *Front. Chem.*, 2022, **10**, 992236.
- 52 H. K. Chin, P. Y. Lin, J. Chen, R. Kirankumar, Z. H. Wen and S. Hsieh, *Appl. Sci.*, 2021, **11**, 4914.
- 53 D. Korcoban, A. E. Kandjani, V. E. Coyle and E. K. Alenezy, *App. Surf. Sci.*, 2022, **578**, 151852.
- 54 U. Malik, D. Korcoban, S. Mehla, A. E. Kandjani, Y. M. Sabri, S. Balendhran and S. K. Bhargava, *App. Surf. Sci.*, 2022, **594**, 153383.
- 55 S. P. Godad and S. Kamal, *Spectrochim. Acta, Part A*, 2021, **236**, 120176.
- 56 H. Qiu, M. Wang, M. Cao, L. Zhang, S. Ji, J. Dou, Y. Ji, S. Kou, J. Guo and Z. Yang, *App. Surf. Sci.*, 2019, **489**, 1010–1018.
- 57 S. Ji, S. Kou, M. Wang, H. Qiu, X. Sun, J. Dou and Z. Yang, *App. Surf. Sci.*, 2019, **489**, 1002–1009.
- 58 H. Qiu, M. Wang, L. Zhang, M. Cao, Z. Yang, J. Dou, S. Ji, Y. Ji, S. Kou and A. S. Bhatti, *Carbon*, 2019, **152**, 305–315.
- 59 W. Pan, L. Liu, Y. Liu, S. Xu, J. Zhou, A. Tang and J. Xue, *J. Environ. Chem. Eng.*, 2023, **11**, 111355.
- 60 G. Su, L. Liu, X. Liu, L. Zhang, J. Xue and A. Tang, *Ceram. Int.*, 2021, **47**, 5374–5387.
- 61 G. Su, L. Liu, Q. Kuang, X. Liu, W. Dong, M. Niu, A. Tang and J. Xue, *J. Mol. Liq.*, 2021, **335**, 116566.
- 62 S. Kumar, R. D. Kaushik and L. P. Purohit, *J. Colloid Interface Sci.*, 2023, **632**, 196–215.
- 63 S. Kumar, R. D. Kaushik and L. P. Purohit, *J. Hazard. Mater.*, 2022, **424**, 127332.
- 64 X. Zhang, C. Shao, X. Li, J. Miao, K. Wang, N. Lu and Y. Liu, *J. Alloy. Compd.*, 2016, **686**, 137–144.

