

Cite this: *Chem. Sci.*, 2021, 12, 6117

All publication charges for this article have been paid for by the Royal Society of Chemistry

# Adaptive response by an electrolyte: resilience to electron losses in a dye-sensitized porous photoanode†

Frances A. Houle \*

Photovoltage and photocurrents below theoretical limits in dye-sensitized photoelectrochemical solar energy conversion systems are usually attributed to electron loss processes such as dye–electron and electrolyte–electron recombination reactions within the porous photoanode. Whether recombination is a major loss mechanism is examined here, using a multiscale reaction–diffusion computational model to evaluate system characteristics. The dye-sensitized solar cell with an  $I^-/I_3^-$  redox couple is chosen as a simple, representative model system because of the extensive information available for it. Two photoanode architectures with dye excitation frequencies spanning 1–25 s<sup>−1</sup> are examined, assuming two distinct recombination mechanisms. The simulation results show that although electrolyte–electron reactions are very efficient, they do not significantly impact photoanode performance within the system as defined. This is because the solution-phase electrolyte chemistry plays a key role in mitigating electron losses through coupled reactions that produce  $I^-$  within the photoanode pores, thereby cycling the electrolyte species without requiring that all electrolyte reduction reactions take place at the more distantly located cathode. This is a functionally adaptive response of the chemistry that may be partly responsible for the great success of this redox couple for dye-sensitized solar cells. The simulation results provide predictions that can be tested experimentally.

Received 21st January 2021  
Accepted 24th March 2021

DOI: 10.1039/d1sc00384d

rsc.li/chemical-science

## Introduction

In photoelectrochemical solar energy conversion systems, which rely on a diffuse energy resource, any factor leading to loss of efficiency in generating electrons or chemical products must be identified and minimized. This is a significant challenge because of the complexity of the chemistry and physics involved. For example, electron losses due to recombination with chemical species in the local environment have been identified as a cause of reduction in photocurrent and photovoltage of dye-sensitized solar cells (DSCs),<sup>1–7</sup> and considerable effort has been made to synthesize structures<sup>8,9</sup> and chemical compositions<sup>10–13</sup> that mitigate them. These elegant approaches often had less effect than hoped, signaling that the mechanisms responsible for performance losses are not fully understood. Photoelectrochemical solar energy conversion systems are highly integrated structures, which makes the design and interpretation of experimental studies that uniquely examine loss processes in them challenging. In this work we take an alternative approach, using a physically-based, detailed multi-scale reaction–diffusion computational model that allows the

inclusion and omission of well-defined mechanistic steps to pinpoint which of them cause losses, and which do not.

DSCs are ideal model systems for this type of study because their function is relatively simple, and they enjoy a detailed yet not fully settled literature on the mechanisms of electron losses. Typical DSCs are constructed of photoanodes that are commonly a porous network of TiO<sub>2</sub> nanoparticles in contact with a conductor, paired with a counterelectrode.<sup>14</sup> The nanoparticle surfaces are coated with dyes that inject electrons into the TiO<sub>2</sub> after photoexcitation, and the network is infiltrated with an electrolyte containing a redox couple such as  $I^-/I_3^-$  that circulates electrons through the system. The chemical reactions taking place in the system are illustrated in Fig. 1.  $I_3^-$  is formed by disproportionation of  $I_2^-$  in the photoanode during operation, and diffuses through bulk electrolyte to a cathode, where it is reduced to  $I^-$ . The  $I^-$  carries electrons back to the photoanode thereby completing the circuit. Transient photoresponse measurements on DSCs have been made to determine electron lifetimes following injection from the dyes, and extract information on electron loss rates in open circuit structures.<sup>15–19</sup> According to these investigations, transient decay is attributed to interfacial processes such as dye–electron recombination and electrolyte–electron recombination in the photoanode.

Of the two recombination channels, dye–electron recombination (sometimes referred to as back electron transfer) has been determined to be relatively unimportant, whereas

Chemical Sciences Division, Lawrence Berkeley National Laboratory, 1 Cyclotron Road, Berkeley, CA 94720, USA. E-mail: fahoule@lbl.gov

† Electronic Supplementary Information (ESI) available. See DOI: 10.1039/d1sc00384d

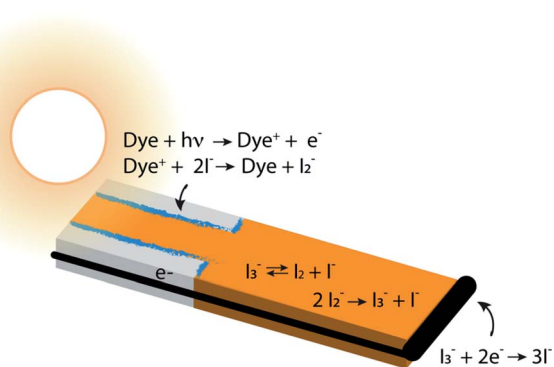


Fig. 1 Cross section through a single pore in a DSC photoanode illuminated from the anode contact side (pore bottom) by sunlight, delineating anodic, bulk electrolyte and cathodic regions and showing the locations of the major chemical reactions. The aspect ratio is not to scale – typical pore diameters are in the range of 10 nm, and the photoanode and overall device thicknesses are usually in the range of 10–20 and 40–80  $\mu\text{m}$ , respectively. The gray regions are  $\text{TiO}_2$  which is transparent to visible light, the blue regions are an adsorbed dye layer, the iodine-containing electrolyte is orange, and the electron conductor and cathode are black. A wire connects the bottom of the anode pore to the cathode, where electrolyte reduction reactions take place. The concentrations of the majority electrolyte ions,  $\text{I}^-$  and  $\text{I}_3^-$ , are connected via an equilibrium reaction. When the DSC is operating, reduction of the oxidized dye,  $\text{Dye}^+$ , leads to a significant local depletion in  $\text{I}^-$  concentration within the pore. The  $\text{I}_2^-$  formed as a result of  $\text{I}^-$  reactions with  $\text{Dye}^+$  undergoes a disproportionation reaction to regenerate  $\text{I}^-$  and  $\text{I}_3^-$ . This leads to a buildup of  $\text{I}_3^-$  in the photoanode pores, forming a concentration gradient that drives its diffusion toward the cathode, where it is reduced to  $3\text{I}^-$ .

electrolyte–electron interactions are thought to be the major loss process.<sup>1,15–17,20–23</sup> Two species present in the  $\text{I}^-/\text{I}_3^-$  electrolyte have been identified as potential reactants with electrons at the photoanode–dye–electrolyte interface. In early studies, it was proposed that  $\text{I}_3^-$  is reduced by trapped electrons, similar to the reaction that occurs at the system's cathode.<sup>7,15</sup> In fact, further investigations have shown that  $\text{I}_2$ , formed in the bulk electrolyte, is the most likely solution-phase species to scavenge photogenerated electrons, generating  $\text{I}_2^-$  as a product.<sup>24,25</sup> However, because the equilibrium constant  $K_{\text{eq}}$  for the reaction  $\text{I}_3^- \rightleftharpoons \text{I}_2 + \text{I}^-$  strongly favors  $\text{I}_3^-$ , the  $\text{I}_2$  concentration has been expected to be low in the presence of  $\text{I}^-$ . Accordingly, it is thought that the recombination rate with  $\text{I}_2$  would be slow, and electrolyte–electron recombination reactions would be likely to be a minor influence on the overall system function.

This suggestion is reasonable, but it carries an underlying assumption that equilibrium among the iodine and iodide species is maintained and overwhelmingly forms  $\text{I}_3^-$  under all conditions.<sup>26</sup> The chemistry of this redox couple is complex,<sup>12</sup> however, and both interfacial and solution phase reactions contribute to the instantaneous local concentrations of iodine-containing species in the electrolyte (Fig. 1). In particular, enrichment of  $\text{I}_3^-$  and depletion of  $\text{I}^-$  in the pores of a working photoanode has been recognized in previous studies.<sup>27,28</sup> This could affect the local concentration of  $\text{I}_2$  (an electron scavenger) via shifting the  $\text{I}^-/\text{I}_3^-$  ratio away from its equilibrium value. Such a shift could provide significant quantities of  $\text{I}_2$  in the

pores, potentially increasing the importance of electron recombination beyond what has been estimated. Establishing experimentally whether or not this occurs is very challenging because of the inaccessibility of nanometer-scale information on the spatially-dependent chemical composition and electron density distributions in nanoporous DSC photoanodes during operation.

Detailed, physically-based multiscale reaction–diffusion calculations offer a means to overcome this limitation and understand in a holistic way how electrolyte chemistry, photo-excitation in the anode, and electron losses are coupled. In the present work, simulations are used to address a simple question: *are electron–electrolyte recombination reactions in the photoanode uniquely responsible for performance limitations of the full DSC system?* The results show that, within the model structure used, the answer is *no*, and that an unexpected mechanism is involved. The calculations provide evidence that homogeneous electrolyte chemistry plays a key role in controlling the system's response to light *operando* and may be part of the reason that the  $\text{I}^-/\text{I}_3^-$  redox couple has been so successful as a component of DSC systems. Specifically, the solution-phase reactions confer a notable resilience on the system by providing redundancy in the generation of  $\text{I}^-$ , a reagent crucial to  $\text{Dye}^+$  reduction and therefore the efficiency of dye regeneration. This is a local, adaptive response of the electrolyte, driven by changing conditions in the photoanode and by the participation of electrolyte–electron recombination reactions. The idea that chemical reactions are central to adaptive behaviors of systems is not new, however reported studies generally examine the adaptivity of molecular structure to external stimuli. In the present work, the chemistry enables *adaptive function* that mitigates what otherwise would be pure losses in the photoanode.

## Model construction

The multiscale, reaction–diffusion framework used in this work has been constructed using experimental data available in the literature, permitting a connection to be drawn between events occurring at the nanometer scale and full system function at the 10 s of microns scale *via* the system's master equation.<sup>29</sup> The framework's basic characteristics, which have been published previously,<sup>28</sup> are summarized. How the starting framework has been extended to include electron recombination reactions and actively maintained equilibria as local concentrations change is described in detail.

All simulations were performed using the kinetics simulation package Kinetiscope,<sup>30</sup> which is a type of kinetic Monte Carlo calculation that uses particles to represent molecular species, and generates a solution to the master equation for the system using stochastic techniques.<sup>31,32</sup> This package builds on algorithms originally described for single compartments<sup>33,34</sup> to support full 3-D reaction–diffusion kinetics with local spatial resolution. The simulations span nm–micron as appropriate, and generate a complete time history for spatially-resolved concentrations of all species in the system. The particle-based simulation technique permits use of non-chemical species, denoted as marker species, to record the occurrence of specific



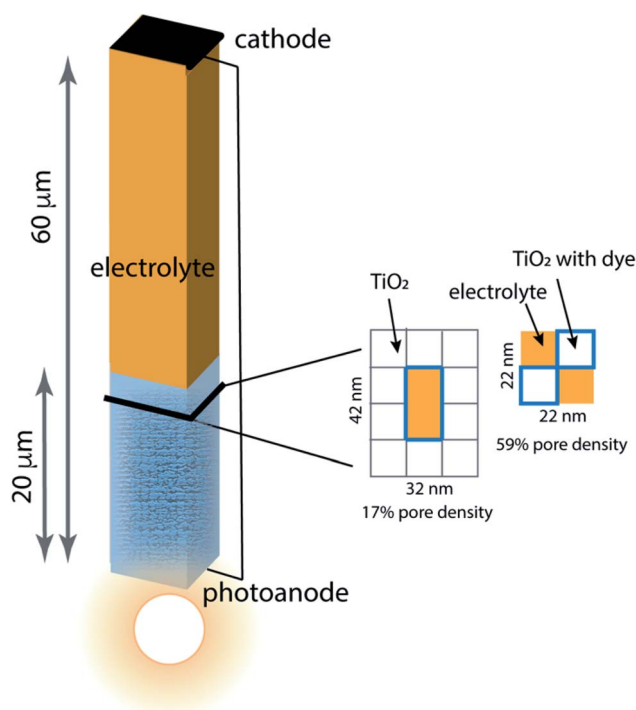


Fig. 2 Schematic of the model architecture using the color scheme from Fig. 1. The photoanode is represented using a single pore surrounded by stacks of  $10\text{ nm} \times 10\text{ nm} \times 10\text{ nm}$   $\text{TiO}_2$  nanoparticles that are coated with dye, filled with electrolyte and in contact with bulk electrolyte. Two limiting pore architectures are considered, and their cross-sectional structure is shown.

types of reactions in specific locations. Recent examples of how this computational technique aids in analysis of the simulation results are presented elsewhere,<sup>28,35,36</sup> details pertinent to how it is implemented in this work are presented in several sections of the ESI.<sup>†</sup>

### Base model framework

Porous photoanodes in DSCs are constructed of  $\text{TiO}_2$  nanoparticles arranged in a complex, randomly packed structure. In previous work,<sup>28</sup> two limiting model architectures were defined to represent this structure in a simplified way that permits connection between function and geometry to be made. Briefly, the system is represented by a single pore of well-defined length and cross section, as illustrated in Fig. 2. The interior walls of the pore are covered with a generic photosensitizing dye that has typical light absorption and excitation properties. Because a nanoporous photoanode will have a broad range of pore sizes and proximities due to the random nature of nanoparticle packing, two limiting pore density cases are considered: 17% pore density, and 59% pore density. In both cases the pores are  $20\text{ }\mu\text{m}$  long, filled with an  $\text{I}^-/\text{I}_3^-$  electrolyte, and separated from the cathode by a  $40\text{ }\mu\text{m}$  thick layer of bulk electrolyte. The mnemonic used to name them is Pore Density  $x\%$ –pore length, for example the 59% pore density– $20\text{ }\mu\text{m}$  pore length architecture is called PD59-20. The anode contact at the bottom of the pore is connected by a zero-resistance wire to the cathode contact above the electrolyte layer. The  $\text{TiO}_2$  anode material is modeled as  $10\text{ nm}$  particles in linear stacks surrounding the pores. The 17% density case has a  $10 \times 20\text{ nm}$  cross sectional area fully surrounded by  $\text{TiO}_2$ , and the 59% density case is a  $10 \times 10\text{ nm}$  cross

Table 1 List of reaction steps and kinetic parameters used in the simulations

| Process                                   | Step                                                                                                                                                                                        | Rate constant                                                                                                                                                                                                                                                                                 |
|-------------------------------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| 1. Equilibrium                            | $\text{I}_3^- \rightarrow \text{I}_2 + \text{I}^-$<br>$\text{I}_2 + \text{I}^- \rightarrow \text{I}_3^-$                                                                                    | $3 \times 10^3\text{ s}^{-1a}$<br>$5 \times 10^{-12}\text{ cm}^3\text{ mol}^{-1}\text{ s}^{-1b}$                                                                                                                                                                                              |
| 2. Anode–electrolyte interface reaction   | $\text{Dye} + h\nu \rightarrow \text{Dye}^+ + \text{electron}$<br>$\text{Dye}^+ + \text{I}^- \rightarrow \text{Dye-I}$<br>$\text{Dye-I} + \text{I}^- \rightarrow \text{Dye} + \text{I}_2^-$ | $1\text{--}25\text{ s}^{-1}\text{ (ref. 14)}^c$<br>$1 \times 10^{-10}\text{ cm}^3\text{ mol}^{-1}\text{ s}^{-1d}$<br>$1 \times 10^{-10}\text{ cm}^3\text{ mol}^{-1}\text{ s}^{-1d}$                                                                                                           |
| 3. Electrolyte reaction                   | $2\text{I}_2^- \rightarrow \text{I}^- + \text{I}_3^-$                                                                                                                                       | $3.8 \times 10^{-11}\text{ cm}^3\text{ mol}^{-1}\text{ s}^{-1}\text{ (ref. 12)}$                                                                                                                                                                                                              |
| 4. Cathode–electrolyte interface reaction | $\text{I}_3^- + 2\text{ electrons} \rightarrow 3\text{I}^-$                                                                                                                                 | $1 \times 10^{-20}\text{ cm}^6\text{ mol}^{-2}\text{ s}^{-1d}$                                                                                                                                                                                                                                |
| 5. Interfacial electron recombination     | $\text{I}_2 + \text{e}^- \rightarrow \text{I}_2^-$                                                                                                                                          | $6.64 \times 10^{-18}\text{ cm}^3\text{ mol}^{-1}\text{ s}^{-1}\text{ (ref. 25)}$                                                                                                                                                                                                             |
| 6. Interfacial $\text{I}_2$ reaction      | $\text{I}_2 + \text{I}^- \rightarrow \text{I}_3^-$                                                                                                                                          | $5 \times 10^{-12}\text{ cm}^3\text{ mol}^{-1}\text{ s}^{-1b}$                                                                                                                                                                                                                                |
| 7. Electron diffusion                     | Electrolyte interface to anode $\text{TiO}_2$ matrix<br>In $\text{TiO}_2$ matrix toward anode contact<br>From anode contact to cathode contact                                              | $2 \times 10^{-4}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 14)}^e$<br>$2 \times 10^{-4}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 14)}^e$<br>$1\text{ cm}^2\text{ s}^{-1a}$                                                                                                                          |
| 8. Electrolyte ion diffusion              | From cathode to electrolyte interface<br>$\text{I}_3^-$ in electrolyte<br>$\text{I}_2^-$ in electrolyte<br>$\text{I}^-$ in electrolyte                                                      | $2 \times 10^{-4}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 14)}^e$<br>$1 \times 10^{-5}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 37 and 38)}^e$<br>$1 \times 10^{-5}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 37 and 38)}^e$<br>$2 \times 10^{-5}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 37 and 38)}^e$ |
| 9. Ion adsorption at interface            | $\text{I}^-$ to anode–electrolyte interface                                                                                                                                                 | $2 \times 10^{-5}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 37 and 38)}^e$                                                                                                                                                                                                                        |
| 10. Ion desorption from interface         | $\text{I}_3^-$ from anode–electrolyte interface<br>$\text{I}_2^-$ from anode–electrolyte interface                                                                                          | $1 \times 10^{-5}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 37 and 38)}^e$<br>$1 \times 10^{-5}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 37 and 38)}^e$                                                                                                                                              |
| 11. Electrolyte $\text{I}_2$ diffusion    | $\text{I}_2$ in electrolyte                                                                                                                                                                 | $1 \times 10^{-5}\text{ cm}^2\text{ s}^{-1}\text{ (ref. 37 and 38)}^e$                                                                                                                                                                                                                        |

<sup>a</sup> Calculated from measured  $K_{\text{eq}}$  and Smoluchowski equation for the reverse step. <sup>b</sup> Calculated from the Smoluchowski equation. <sup>c</sup> This represents a complex sequence of excitations, energy transitions, relaxations and charge injection, all taking place on the ps–μs timescale. The overall frequency of this step is in the range of fractions of seconds, however, because of the low light intensity. <sup>d</sup> Assumed value for rate constant, selected to be very large. <sup>e</sup> Estimated from range in cited references.

sectional checkerboard with wraparound diffusion paths to make a boundaryless system. Neglect of tortuosity is not likely to significantly affect the calculations since its main influence would be to slow down overall diffusion rates within the anode. The dye sensitizer molecules are assumed to be in a uniform coating 1 nm thick on all surfaces in contact with the electrolyte, with a concentration of  $4.4 \times 10^{20} \text{ mol cm}^{-3}$ .<sup>14</sup> The electrolyte composition has an initial value of  $3.01 \times 10^{19} \text{ mol cm}^{-3}$   $\text{I}_3^-$  and  $3.01 \times 10^{20} \text{ mol cm}^{-3}$   $\text{I}^-$  throughout the system, and diffusion and reaction coefficients for an acetonitrile solution. Electrons injected into the  $\text{TiO}_2$  after photon absorption by the dye diffuse to an anode contact, and are conducted to the cathode where reduction of  $\text{I}_3^-$  to form  $\text{I}^-$  takes place. Diffusion of the  $\text{I}^-$  back to the anode pore serves to complete the circuit. The reaction steps and rate constants used in this work build on the base model, and are presented in Table 1. The base model is processes 2–4 and 7–10. Processes 1, 5, 6, and 11 are new in this work.

### Modeling electron recombination with $\text{I}_2$

In order to examine the influence of electron recombination with  $\text{I}_2$ , a source for it has to be explicitly included in the electrolyte chemistry. The initial concentration of  $\text{I}_3^-$  relative to  $\text{I}^-$  and  $\text{I}_2$  is governed by the equilibrium constant  $K_{\text{eq}}$  in acetonitrile,  $1 \times 10^6 \text{ M}^{-1}$ .<sup>25,39</sup> This value is at the low end of the range in the literature,  $10^6$ – $10^7 \text{ M}^{-1}$ ,<sup>12,25,39–41</sup> and consistent with the analysis that generated the  $\text{I}_2$  recombination rate coefficient.<sup>25</sup> (The sensitivity to using this low value is not significant, as discussed in the ESI, Section 2.†) When the electrolyte is formulated, the  $\text{I}_2$  concentration is extremely low because it is added to an excess of  $\text{I}^-$  and is converted essentially entirely to  $\text{I}_3^-$ . If the initial  $\text{I}^-$  and  $\text{I}_3^-$  ion concentrations are unaltered during system operation, equilibrium is maintained and the  $\text{I}_2$  concentration should remain low. However, during operation the  $\text{I}^-/\text{I}_3^-$  ratio decreases in the pores,<sup>27,28</sup> which can increase the  $\text{I}_2$  concentration according to

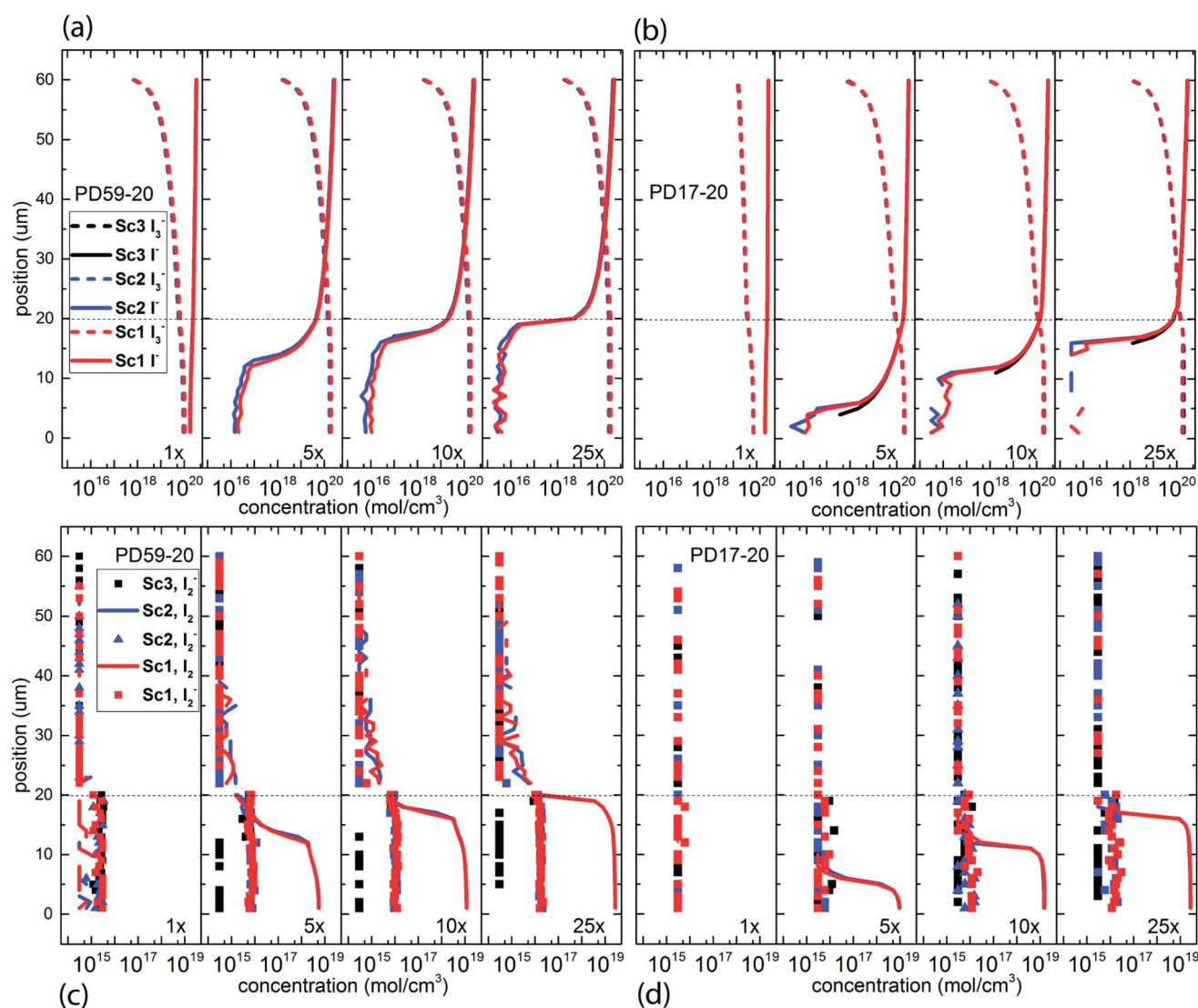


Fig. 3 Electrolyte composition at 0.8 s (close to steady state) as a function of position within the system. The photoanode contact is at 0  $\mu\text{m}$ , the interface between the photoanode and bulk electrolyte is at 20  $\mu\text{m}$ , and the cathode is at 60  $\mu\text{m}$ . (a)  $\text{I}^-$  and  $\text{I}_3^-$  profiles for PD59-20, (b)  $\text{I}^-$  and  $\text{I}_3^-$  profiles for PD17-20, (c)  $\text{I}_2^-$  and  $\text{I}_2$  profiles for PD59-20, (d)  $\text{I}_2^-$  and  $\text{I}_2$  profiles for PD17-20.



the reversible reactions, shown in Table 1, process 1. Rate constants for these forward and reverse steps in acetonitrile have not been reported. The Smoluchowski equation is used to estimate the rate constant for  $I^- + I_2 \rightarrow I_3^-$  assuming van der Waals radii of 0.2 nm for the reactants, and a diffusion coefficient of  $10^5 \text{ cm}^2 \text{ s}^{-1}$  (Table 1). The value of  $5 \times 10^{-12} \text{ cm}^2 \text{ s}^{-1}$  is about half of the value measured in protic solution.<sup>42</sup> The dissociation rate constant is calculated using it, and  $K_{\text{eq}}$ .  $I_2$  is assumed to diffuse freely throughout the electrolyte, and from the electrolyte into the interfacial region where the dye is located (Table 1, process 11). Electrons are initially injected by excited dyes in this region, and are assumed to be available for recombination with electrolyte species when they are present at the interface. Electrons can diffuse into the  $\text{TiO}_2$  bulk with a rate determined by the electron concentration gradient between the interface and the bulk and their diffusion coefficient. Once in the bulk, they are no longer available for recombination with  $I_2$ . The energetic state of the electrons (conduction band vs. mid-gap states) is important,<sup>4,43</sup> but not considered in this model because insufficient information is available about the relative concentrations of these states and their reactivity toward  $I_2$  in the interfacial region. The

appropriate kinetics can be added readily as experimental or theoretical data are published.

The details of how the electron recombination with  $I_2$  should be described are not fully known,<sup>4</sup> so three scenarios are examined in this study. Scenario 1 (Sc1) assumes that electron transfer to  $I_2$  to form  $I_2^-$  is an elementary step, with the measured rate constant of  $6.64 \times 10^{-18} \text{ cm}^3 \text{ mol}^{-1} \text{ s}^{-1}$ ,<sup>25</sup> which is significantly slower than the values of  $10^{-15}$ – $10^{-16} \text{ cm}^3 \text{ mol}^{-1} \text{ s}^{-1}$  estimated from full system models.<sup>27</sup> This scenario is represented as process 5 in Table 1. Scenario 2 (Sc2) assumes that process 5 competes with reaction between  $I^-$  and  $I_2$  at the interface (process 6), which could potentially reduce the extent of electron recombination. The rate coefficient is taken to be the same value as in solution (process 1). As a control, scenario 3 (Sc3) assumes no  $I_2$  formation, which corresponds to the previously published base case<sup>28</sup> described above.

### Performance scaling with increasing excitation

The base case from earlier work used a dye excitation frequency of  $1 \text{ s}^{-1}$  (process 2, Table 1), which is typical for 1 sun.<sup>15,44</sup>

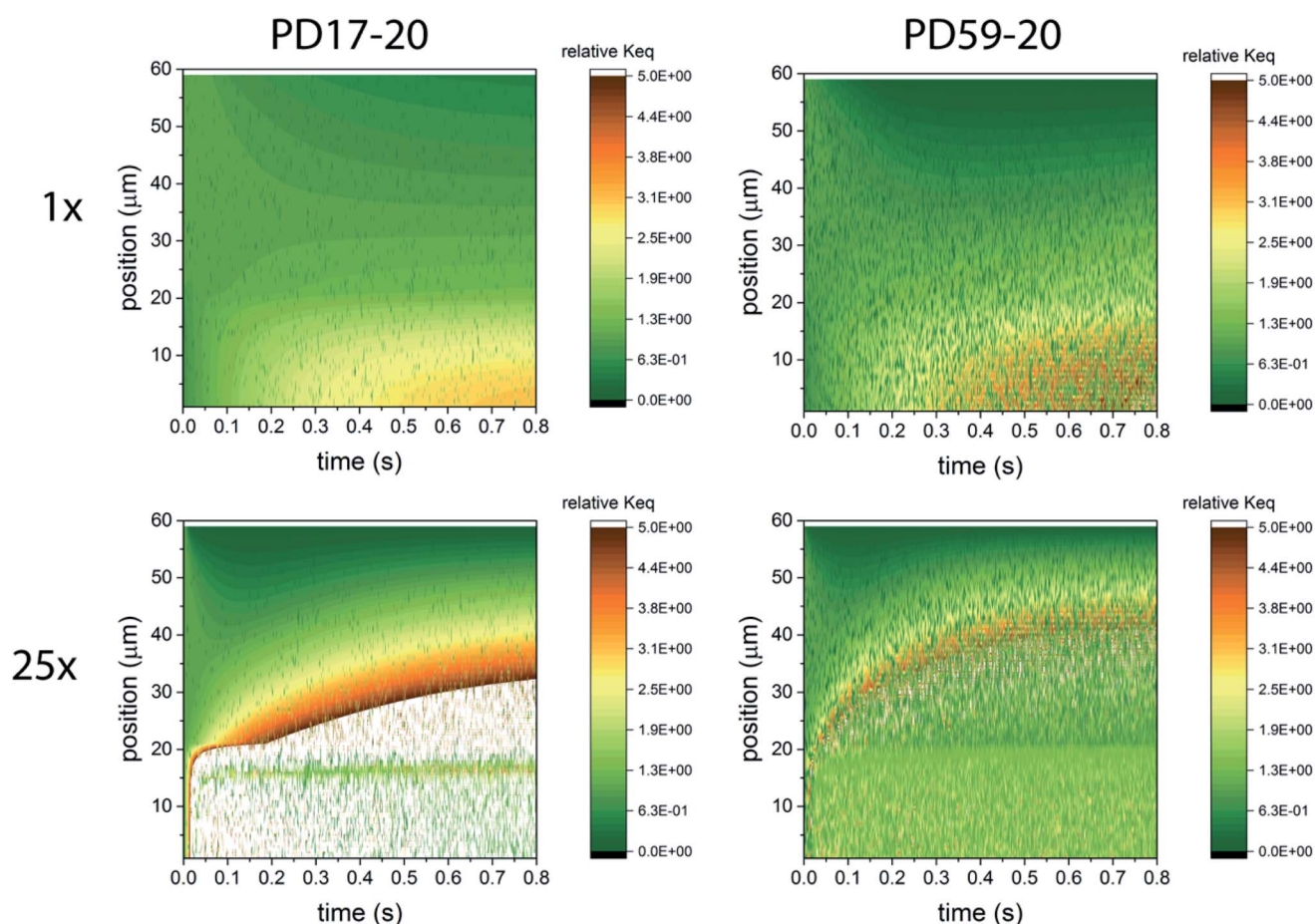


Fig. 4 Maps of the ratio  $([I_3^-]/[I^-][I_2])/K_{\text{eq}}$  as a function of position and time in the electrolyte, as a measure of deviation from expected equilibrium concentrations under Sc2. The photoanode contact is at 0  $\mu\text{m}$ , the interface between the photoanode and bulk electrolyte is at 20  $\mu\text{m}$ , and the cathode is at 60  $\mu\text{m}$ . In the simulations 1 particle represents  $3 \times 10^{14} \text{ mol cm}^{-3}$ . Whenever the ion concentration is 0 particles, it is set to a value of  $6 \times 10^{13} \text{ mol cm}^{-3}$  to permit these ratios to be calculated. (a)  $1 \text{ s}^{-1}$  excitation frequency, PD17-20, (b)  $25 \text{ s}^{-1}$  excitation frequency, PD17-20, (c)  $1 \text{ s}^{-1}$  excitation frequency, PD59-20, (d)  $25 \text{ s}^{-1}$  excitation frequency, PD59-20. Data for Sc1 are nearly indistinguishable from this set.

Higher levels of excitation leading to higher rates of electron generation and recombination have been found to impact photoanode function.<sup>11,27,45</sup> Accordingly, the 1 sun case is compared to 5 $\times$ , 10 $\times$  and 25 $\times$  higher excitation levels for all 3 scenarios. These are represented in the model using the dye excitation step in process 2, Table 1, with rate constants of 5 s<sup>-1</sup>, 10 s<sup>-1</sup> and 25 s<sup>-1</sup>. Physically, this would correspond to using more efficient dyes or use of concentrated light, or both.

## Results

Simulations were performed for both DSC architectures (Fig. 2) with 4 excitation frequencies and 3 recombination scenarios, a total of 24 data sets. Each simulation was allowed to run to 0.8 s, corresponding to steady state for nearly all runs. This time is relatively short, and is determined by the establishment of steady state gradients in the system which has very thin (1 micron) electrodes. Charging up thicker, more realistic electrode substrates is too expensive computationally, but would yield a more realistic time to reach steady state and potentially result in lower average electron densities in the system.

The simulation outputs provide time histories for I<sub>3</sub><sup>-</sup>, I<sup>-</sup>, I<sub>2</sub>, electrons, and neutral and oxidized dyes for all systems. In addition, spatiotemporal maps for marker species provide insights to interactions and events that underly the composition *vs.* time results. Referring to Table 1, markers track each occurrence of I<sub>2</sub> formation (process 1) and electron recombination by I<sub>2</sub> (process 5) in Sc1 and Sc2, interfacial I<sub>3</sub><sup>-</sup> formation (process 6), in Sc2, and disproportionation (process 3), excitation and Dye<sup>+</sup> reduction (process 2), and I<sub>3</sub><sup>-</sup> reduction at the cathode (process 4), in all 3 scenarios. The full set of simulation results including the markers is analyzed to understand how each scenario affects system performance.

### Effect of architecture, excitation and recombination on electrolyte composition

The electrolyte composition through the full thickness of both architectures at 0.8 s is shown in Fig. 3, and full maps are shown in the ESI, Fig. S2–S9.† The I<sup>-</sup> and I<sub>3</sub><sup>-</sup> profiles in Fig. 3a and b are nearly the same for all 3 scenarios for each architecture, showing accumulation of I<sub>3</sub><sup>-</sup> and deep depletion of I<sup>-</sup> in the anode pores, and depletion of I<sub>3</sub><sup>-</sup> near the cathode. The concentration polarization is modest at 1 s<sup>-1</sup> excitation frequency, and increases significantly with increasing excitation frequency. I<sup>-</sup> is more deeply depleted in PD59-20 because of the larger surface/volume ratio in the pores.<sup>28</sup> Fig. 3c and d show the corresponding I<sub>2</sub> and I<sub>2</sub><sup>-</sup> populations for all 3 scenarios. It is evident that as the I<sup>-</sup> concentration is reduced, I<sub>3</sub><sup>-</sup> is driven toward I<sub>2</sub> and I<sub>2</sub><sup>-</sup> formation. I<sub>2</sub><sup>-</sup> is formed by either I<sub>2</sub> recombination with electrons from TiO<sub>2</sub> (scenarios 1 and 2) or by reaction of I<sup>-</sup> with oxidized dyes at the anode–electrolyte interface (all scenarios). I<sub>2</sub><sup>-</sup> undergoes a second order disproportionation reaction (Table 1, process 3) to regenerate I<sub>3</sub><sup>-</sup> and I<sup>-</sup> within the pores. As shown in ESI, Fig. S10,† the total number of disproportionation events is significantly increased when electron recombination occurs, constituting a local source of I<sup>-</sup>

in addition to the cathodic reactions. The I<sub>3</sub><sup>-</sup> formed in process 3 can dissociate to form still more I<sub>2</sub> and I<sup>-</sup> within the pore.

In Sc3, where electron recombination does not occur, I<sub>2</sub><sup>-</sup> is formed only in regions where I<sup>-</sup> is present to react with oxidized dyes on the pore walls. Since I<sup>-</sup> is depleted in the pore and transport from the cathode is not fast enough to replace it, I<sub>2</sub><sup>-</sup> formation and disproportionation become progressively confined to the top of the photoanode (ESI, Fig. S10†). Under Sc1 and Sc2, on the other hand, disproportionation is spread throughout the pore volume. The concentrations of both I<sub>3</sub><sup>-</sup> and I<sup>-</sup> in the electrolyte are reduced under Sc2, relative to Sc1. In the PD59-20 architecture, 25% of the I<sub>3</sub><sup>-</sup> is located at the TiO<sub>2</sub>–electrolyte interface where it is formed by the interfacial I<sub>2</sub> reaction, Table 1 process 6. In the PD17-20 architecture, 15% of the I<sub>3</sub><sup>-</sup> is at the interface due to the smaller surface/volume ratio. In both cases, I<sub>3</sub><sup>-</sup> accumulates at the interface after it is formed because its diffusion into the bulk electrolyte is gradient-driven and therefore slow.

In considering the relative concentrations of I<sub>3</sub><sup>-</sup>, I<sup>-</sup> and I<sub>2</sub> in an operating DSC, the question arises whether their balance is controlled only by their rapidly maintained equilibrium in solution as would be the case for a buffer. This would be expected if the kinetics allow the electrolyte to respond on

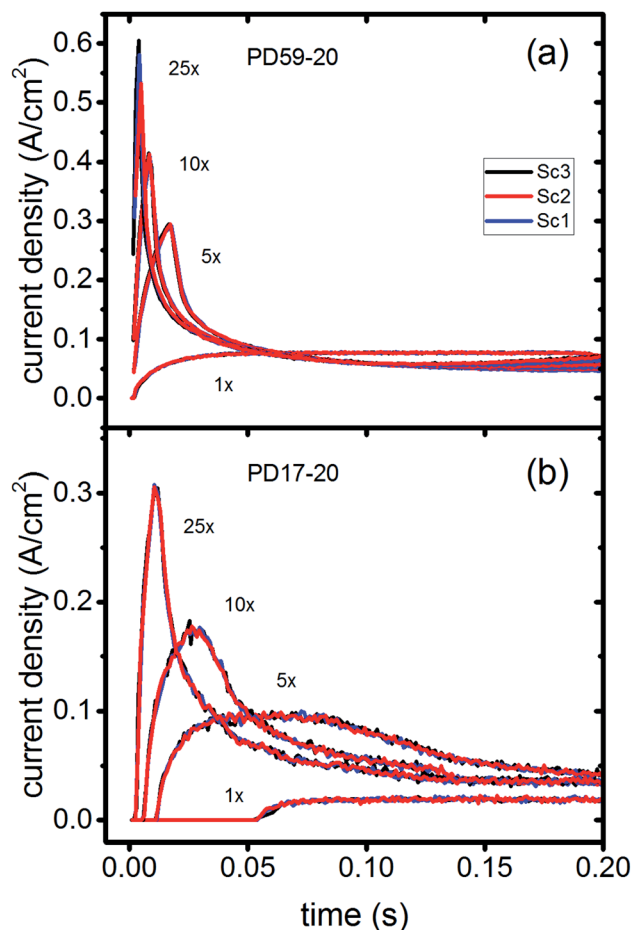


Fig. 5 Current densities for all excitation frequencies to 0.2 s (a) PD59-20; (b) PD17-20.



a timescale that is fast relative to the interfacial reactions. To answer this question, the instantaneous concentration ratio  $[I_3^-]/[I^-][I_2]$  was normalized to  $K_{eq}$  as a function of position and time in the PD17-20 and PD 59-20 systems for the  $1\text{ s}^{-1}$  and  $25\text{ s}^{-1}$  excitation frequency cases, Sc2, as shown in Fig. 4. If equilibrium is maintained and controls all relative concentrations in the electrolyte, then the normalized value should be close to 1. If it is much less than 1, then the system is locally depleted in  $I_3^-$ ; if it is much greater than 1, the system is locally depleted in  $I_2$  and/or  $I^-$ . It is clear that the equilibrium reaction does not consistently control relative concentrations except possibly for within the pore of the PD59-20 architecture at a  $25\text{ s}^{-1}$  excitation frequency. Rather, the electrolyte is driven away from equilibrium and it is the combination of the interfacial and solution phase reactions that controls the electrolyte composition during DSC operation.

### Photocurrent and photovoltage with and without electron recombination

In DSCs, current flow is the result of the reduction of  $I_3^-$  that is formed in the photoanode.  $I_3^-$  diffuses to the cathode to form  $3I^-$ , which then diffuse back to the photoanode. Predicted current densities are calculated by taking the derivative of the accumulating  $I_3^-$  reduction event markers vs. time.<sup>28</sup> Fig. 5 presents the calculated current densities for the three scenarios and both architectures, for all excitation frequencies. For clarity only data to 0.2 s are shown, data to 0.8 s are in ESI, Fig. S11.† Photocurrent transients are predicted for all cases except PD17-20 at  $1\text{ s}^{-1}$  excitation frequency.

The current densities predicted for all three scenarios are similar up to 0.15 s for both architectures. The steady state/near steady state current densities for all cases calculated at 0.8 s are plotted in Fig. 6a, showing that they are similar in value for all 3 scenarios for both pore architectures. Although Sc1 and Sc3 have nearly identical current densities, that for Sc2 is reduced. This is a consequence of allowing adsorption of  $I_3^-$  on the pore walls in scenario 2, which reduces  $[I_3^-]$  in the electrolyte. The surface to volume ratio is larger for PD59-20, so the effect is more pronounced.<sup>28</sup> Both Fig. 5 and 6a show that the highly efficient electron recombination reactions and  $I_2$  reactions with  $I^-$  at the photoanode–electrolyte interface do not influence the photocurrent, and Fig. 6b shows that in all cases photocurrent scales the same way with excitation frequency. This is a surprising result: from the literature, electron recombination is expected to reduce photocurrent, and further addition of an  $I_2$ – $I^-$  recombination path that competes with electron recombination by  $I_2$  might have been expected to be neutral or perhaps increase current density.

Fig. 6c shows the calculated electron densities at 0.8 s in the photoanode. The electron densities are essentially the same for all 3 scenarios in the PD17-20 architecture, while electrolyte–electron recombination reduces them by about 15% in the PD59-20 architecture (Sc1 and Sc2). The electron densities can be used to calculate the relative photovoltage  $\Delta V$  to evaluate to what extent recombination affects it. An absolute open or closed circuit photovoltage requires knowledge of the dark electron population, which the literature indicates can vary over many orders of magnitude.<sup>18,44,46</sup> There is no accepted typical value to

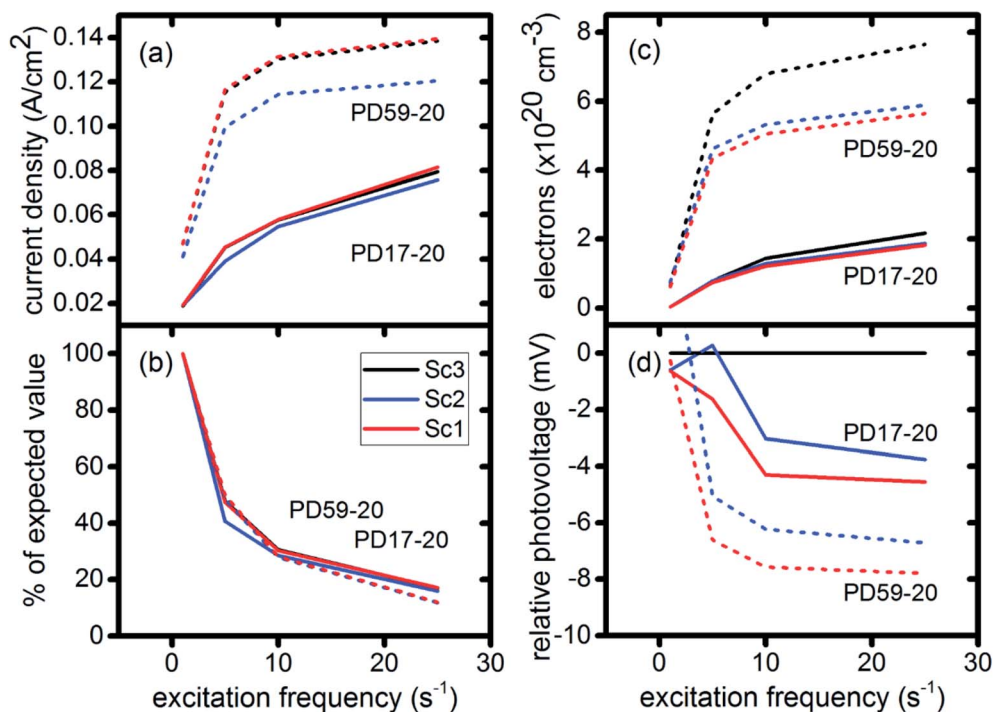


Fig. 6 Electrical performance characteristics for both architectures and all three scenarios as a function of excitation frequency. (a) Calculated current densities at 0.8 s; (b) percent of expected current density values if they had scaled linearly with excitation frequency; (c) electron densities in the  $\text{TiO}_2$ , averaged from top to bottom of the photoanode; (d) photovoltage relative to Sc3 (no electron recombination).



assume for the present work. The relative photovoltage,  $\Delta V^{ab}$ , can be estimated, however, from the change in electron density  $D_e$  in state b relative to that in state a. It is expressed as:

$$\Delta V^{ab} = \frac{e}{kT} \ln \frac{D_e^b}{D_e^a} \quad (1)$$

where  $T$  is assumed to be 298 K and here state a is Sc3 (no electrolyte–electron recombination) and state b is Sc1 or Sc2, at each excitation frequency. Fig. 6d shows that the relative photovoltage decreases between  $1\times$  and  $5\times$  excitation frequencies, but then is essentially flat. Typical measured open circuit potentials are in the range of 600 mV. For both architectures, electron recombination under both scenario 1 and scenario 2 results in a small loss in relative potential compared to the no-recombination case, scenario 3. However, the current densities do not follow this trend, indicating that the overall resistance of the system has decreased as a result of including electron recombination.

### Interfacial chemistry

The interfacial electron recombination and  $I_2$  recombination reaction frequencies mirror the electrolyte composition as

a function of position in the photoanode pores, shown in Fig. 3. The abundance of  $I_2$  present in the electrolyte provides a means for these reactions to become dominant events in the photoanodes. As shown in Fig. 7a, recombination under Sc1 and Sc2 consumes most of the electrons generated by photoexcitation at  $25\text{ s}^{-1}$  excitation frequency, with an architecture-dependent spatial distribution. Fig. 7b shows that under Sc2, recombination of  $I_2$  and  $I^-$  to form  $I_3^-$  dominates over electron recombination at all excitation frequencies.

Photoexcited dyes are the source of electrons that undergo recombination with the electrolyte. Following charge injection, the dyes remain in their inactive, oxidized state ( $\text{Dye}^+$ ) until they react with  $2I^-$  to form  $I_2^-$  (Table 1, process 2). As shown in Fig. 8 and ESI, Fig. S2–S9,† under Sc3 there is a significant persistent population of  $\text{Dye}^+$  in the photoanode for excitation frequencies above  $1\text{ s}^{-1}$ . Because the oxidized dyes cannot undergo further excitation and charge injection, this has the effect of reducing the relative photosensitivity of the anode from 1 at  $1\times$  excitation frequency to 0.97, 0.69 and 0.37 at  $5\times$ ,  $10\times$  and  $25\times$  excitation frequencies for PD17-20, and 0.74, 0.43 and 0.19 across the same series for PD59-20. On the other hand, under Sc1 and Sc2, the fraction of dyes that are active is high, and the relative photosensitivity stays at 1.

Significant current transients are evident in Fig. 5 that cannot be explained by the electrolyte and interfacial simulation results described so far. The full composition vs. time maps for the systems presented in ESI, Fig. S2–S9,† show that no transient concentration changes occur in the photoanode under any conditions. This is not the case for the bulk electrolyte near the cathode, however, where transients in the  $I_3^-$  and  $I^-$  concentrations near the cathode are evident. Profiles for PD17-20 and PD59-20 using Sc2 are shown in Fig. 9. The current density transients for  $25\times$  excitation frequency at around 150

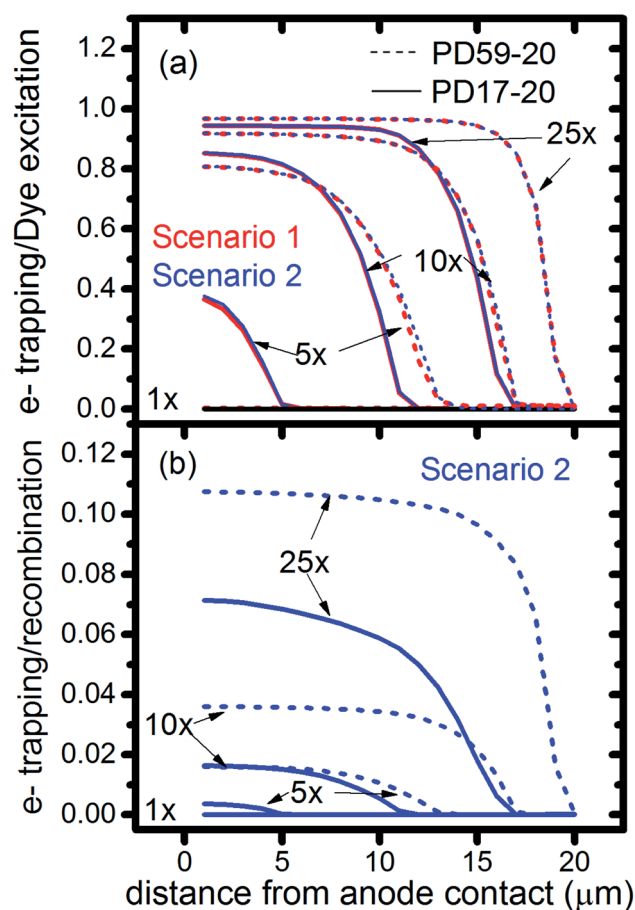


Fig. 7 Ratios of cumulative events as a function of position in the photoanode after 0.8 s, as monitored by markers in the reaction steps. (a) Electron recombination with  $I_2$  relative to dye excitation for Sc1 and Sc2, (b) electron recombination with  $I_2$  relative to interfacial formation of  $I_3^-$  from  $I_2$  and  $I^-$ .

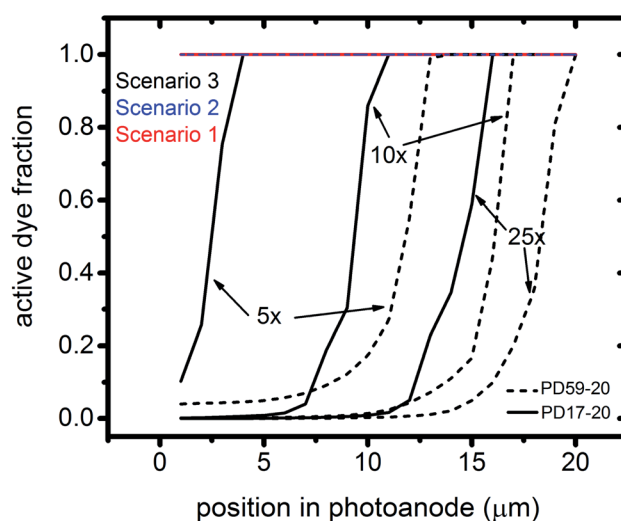


Fig. 8 Profiles of the fraction of dyes that are photoactive within the pores for Sc1 and Sc2, which allow electrolyte–electron recombination, and Sc3, which does not, as a function of excitation frequency. The active dye fraction at  $1\times$  excitation frequency is 1.0 everywhere in the photoanode for both architectures and all 3 scenarios.

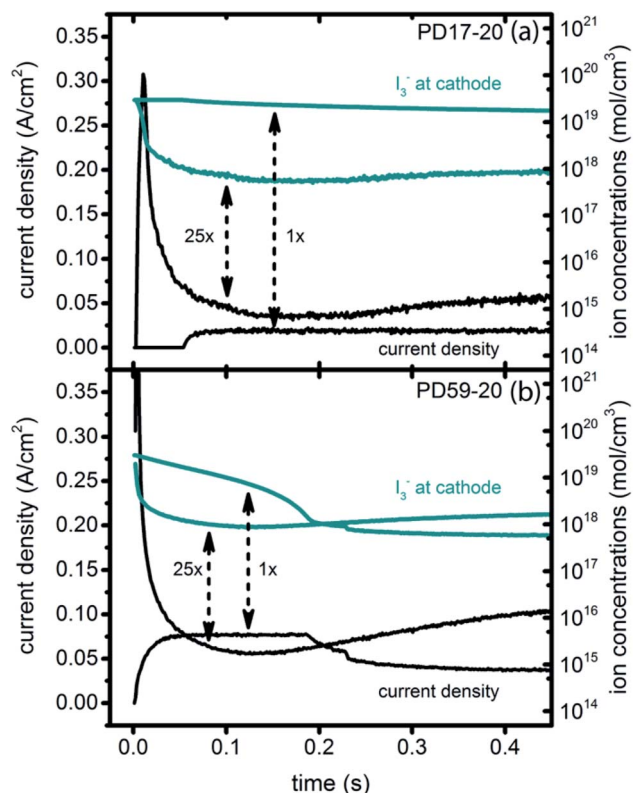


Fig. 9 Current transients (left axis) and  $I_3^-$  concentrations at the cathode (right axis) for Sc2. (a) PD17-20, 1 $\times$  and 25 $\times$  excitation frequencies; (b) PD59-20, 1 $\times$  and 25 $\times$  excitation frequencies.

ms for both cases are accompanied by a depletion in local  $[I_3^-]$ . At 1 $\times$  excitation frequency, the lower pore density case shows no transient and no  $I_3^-$  depletion, and the high pore density case shows a more gradual decline in current density that tracks the decay in  $[I_3^-]$ . Further evidence that cathodic processes control current density is provided by the linear correlation between current density and  $[I_3^-]$  under all conditions except the PD17-20 architecture at 1 $\times$  excitation frequency, as shown in ESI, Fig. S12.†

## Discussion

### Performance losses in DSCs

The primary goal of this work is to understand whether back electron transfer to electrolyte species is a source of performance losses in a simple photoelectrochemical energy conversion system, the DSC. The calculations show that electron losses due to recombination with  $I_2$ , while large, do not significantly reduce photovoltage or photocurrent. The only factor found to do so that directly involves electrolyte in the photoanode is creation of a sink for  $I_3^-$  on pore walls within photoanode, which reduces both the in-pore and cathodic sources of  $I^-$ . The magnitudes of both relative photovoltages and photocurrents do have significant dependences on other characteristics of the system, however. They are dependent on pore geometry, with the 59% porosity case providing higher performance than the

17% case. They also depend on excitation frequency. The photocurrent ratio for the two geometries is 1 : 4 at 1 $\times$  excitation frequency, as shown in the earlier study,<sup>28</sup> and decreases to 1 : 2 above 5 $\times$  excitation frequency. This trend occurs because the photocurrent is influenced in different ways by the photoanode (excitation) and cathode (consumption of electrons). The interplay between them is excitation frequency and transport-dependent. The cause of lower than theoretical performance, often attributed to electron losses alone, must have multiple physical origins.

The calculations provide information on how electrolyte–electron recombination influences DSC operation. It is a dominant process, but it does not affect the predicted current density (aside from the PD59-20, Sc2 case, where  $I_3^-$  is depleted), and only modestly reduces the photovoltage. Current density is strongly influenced by cathodic reactions, however, which depend on the local concentration of  $I_3^-$ . Fig. 3 shows that this concentration is reduced near the cathode relative to its value in the bulk and in the pores for all 3 scenarios at all excitation frequencies. This is an incomplete picture, however, because the efficiency of the photoanode, where electrons are generated, depends on whether or not electrons are trapped by  $I_2$ . As shown in Fig. 8, there is reduction in active dyes when no electrolyte–electron recombination occurs (Sc3), which acts to shut down the photoanode and suppress electron generation. It is possible that a significant population of oxidized dyes could make back electron transfer to  $Dye^+$ , ordinarily very slow, kinetically important. If this occurs, for example by coulombic trapping or a hopping mechanism,<sup>4,47</sup> it would be beneficial because it would increase dye cycling. However, the interfacial electron concentration, which ranges from  $7 \times 10^{19}$ – $8 \times 10^{20} \text{ cm}^{-3}$  for the cases studied here, is still less than the  $Dye^+$  concentration of about  $10^{21} \text{ cm}^{-3}$  and back electron transfer could not be a route to full activation of the photoanode. The entire photoanode remains active under Sc1 and Sc2, however a large fraction of the electrons are trapped by  $I_2$  and are swept into the electrolyte as  $I_2^-$  rather than diffusing to the cathode. Despite these operational differences, the photocurrent is predicted to be nearly the same for all three scenarios. This indicates that for the conditions examined here, how much electrolyte–electron recombination occurs is not a dominant consideration for photocurrent generation. What is significant are the cathodic processes involving the electrolyte, which appear to be not just influential, but controlling (ESI, Fig. S12†).

That the details of the  $I^-/I_3^-$  electrolyte redox chemistry are central to DSC performance has been noted extensively in the literature. It has been proposed that the high efficiency of this redox couple in a device relative to other systems is due to the slow rate of electron scavenging by  $I_2$ , governed by a small rate coefficient and a low  $I_2$  concentration,<sup>24</sup> or by  $I_3^-$ .<sup>48</sup> Another hypothesis is that  $Dye^+$  reduction by the electrolyte generates  $I_2^-$ , whose disproportionation products,  $I^-$  and  $I_3^-$  either do not react with electrons released at the  $TiO_2$ –dye–electrolyte interface ( $I^-$ ), or do so only very slowly.<sup>39</sup> These proposals rest on the underlying assumption that preferential formation of  $I_3^-$  is favored by the equilibrium reactions in process 1 under all operating conditions. In fact, the calculations show that this is

not the case: the DSC is a driven system where equilibria are not necessarily maintained even when steady-state operating conditions are reached. The continuous cycling of dye excitations, charge injection and reduction of the oxidized dye depletes the electrolyte present in the pores of  $\text{I}^-$ , and builds up  $\text{I}_3^-$  *via*  $\text{I}_2^-$  disproportionation. The low concentration of  $\text{I}^-$  in the photoanode pores promotes dissociation of  $\text{I}_3^-$  into  $\text{I}_2$  and  $\text{I}^-$ , providing a dynamic and substantial source of  $\text{I}_2$  and consequently a high back electron transfer rate, as well as a significant local source of  $\text{I}^-$  that rapidly reduces oxidized dyes to make them available for photoexcitation and charge injection. The present work suggests that this chemistry may be more central to the usefulness of this redox couple.

### The electrolyte as an adaptive element that preserves system function

To gain deeper insights to the role of the electrolyte in the operation of this system, it is instructive to consider Fig. 3 and 6–8. Whether or not back electron transfer occurs, both previous work<sup>28</sup> and the present study indicate that the photocurrent depends primarily on cathodic reactions, dye amounts in the pores, dye excitation frequency, and pore architecture. Electron losses are large, as shown in Fig. 7, but do not displace these controlling factors. Fig. 6 shows that the only signature of electron losses is a reduction of the photovoltage by a few mV. The model as constructed indicates that efficient recombination losses can be seen to be beneficial because they provide a means for all dyes to be active at all times within the photoanode. This compensation is due to changes in the rates of reactions in the electrolyte, which are situational, because of changes in concentrations as illustrated in Fig. 4. The dynamic changes in rates are fundamentally an adaptive response that confers functional resilience on the photoanode.

The coupled chemical processes that enable this adaptivity are illustrated in Fig. 10. The pathways in black are those generally invoked in the literature, and the green pathways are those that are added in this work so that the electrolyte chemistry is described more completely. The key process is that depletion of  $\text{I}^-$  promotes formation of  $\text{I}^-$  within the pores through two reaction networks. One is the  $\text{I}_2^-$  disproportionation step involving  $\text{I}_2^-$  formed during  $\text{Dye}^+$  reduction, the other

is  $\text{I}_3^-$  dissociation in solution. The two networks are coupled by  $\text{I}_2^-$  disproportionation which forms both  $\text{I}^-$  and  $\text{I}_3^-$ . This single point of coupling provides a feedback mechanism that is driven by dye excitation and oxidized dye reduction processes. The faster the dye excitation and charge injection, the faster  $\text{I}^-$  is consumed at the interface to generate  $\text{I}_2^-$ . The more  $\text{I}_2^-$  there is, the faster the disproportionation reaction which has a second order dependence on its concentration. This does not fully replace  $\text{I}^-$ , so leads to an acceleration of the rate of  $\text{I}_3^-$  dissociation to form  $\text{I}_2$ .  $\text{I}_2$  combines with electrons at the interface, forming  $\text{I}_2^-$  and further increasing the disproportionation rate. The dual routes to  $\text{I}^-$  formation allow the system to cycle electrons locally *via* iodide ions, within the pores, at a rate sufficient to maintain photoanode efficiency without requiring transport of  $\text{I}_3^-$  to the cathode. Thus, the electrolyte chemistry allows the photoanode's function to be resilient when driven by increasingly larger excitation frequencies and ever-increasing electrolyte–electron recombination. A process that should result in a net loss in the system's performance does not have this impact because it is not decoupled from a network of compensating chemical reactions.

That this electrolyte reaction mechanism is inherently adaptive is a prediction of this detailed computational study. Computational predictions need to be tested experimentally, and the simulation results point to useful measurements that will clarify their validity within a real system. As can be seen in Fig. 3, measurements of  $\text{I}^-$  and  $\text{I}_3^-$  concentrations even within the pores will not enable the importance of the electron back reaction channel and hence the dual channel for  $\text{I}^-$  formation to be assessed. The presence of  $\text{I}_2^-$ , which is a sensitive probe of the  $\text{I}_2$  reaction channel, has not been detected.<sup>24</sup> Consideration of the redox couple chemistry<sup>12</sup> led to an estimate for the  $\text{I}_2^-$  concentration of 3  $\mu\text{M}$  at 10  $\text{mA cm}^{-2}$  photocurrent. The simulation results show that it is much lower, in the range of  $10^{-8}$  M within the pores at an excitation frequency of  $1 \times$ , likely as a result of the high disproportionation rate within those small spaces. The simulations do predict a buildup of  $\text{I}_2$  within the pores at the highest excitation frequencies, however, which should be detectable if spectroscopies probing that region and not the bulk electrolyte are possible. Spectroscopic observations of the bulk electrolyte could be useful for a system under intense illumination. Any imbalances between  $\text{I}_3^-$  and  $\text{I}^-$  due to the system being driven away from equilibrium as shown in Fig. 4 may be evident if the optical path length is sufficient. A potentially more definitive assessment would be *via* measurements of the presence or absence of oxidized dye in the photoanode *operando* as a function of excitation frequency, photoanode preparation techniques, and electrolyte composition. The simulations predict that  $\text{Dye}^+$  reduction is fast under all conditions if electrolyte–electron recombination is an important channel. If oxidized dyes are detected, they would indicate that the recombination and electrolyte chemical mechanisms described in the literature need to be refined. Finally, the detailed kinetics of the  $\text{I}_3^-$  reduction reaction at the cathode are insufficiently understood. The assumptions in the present work lead to a conclusion that it is this reaction that ultimately limits photocurrent, however that could change if its

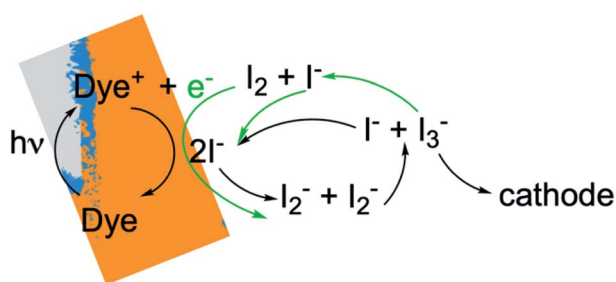


Fig. 10 Schematic of the electrolyte reactions involved in reduction of  $\text{Dye}^+$ . The paths indicated by black arrows correspond to Sc3. The green arrows show how the  $\text{I}_2^-$ –electron recombination reactions from Sc1 and Sc2 are coupled to them.



mechanism proves to be significantly different than as described here.

## Conclusions

According to the literature, photocurrent and photovoltage in DSCs are limited by trapping of photogenerated electrons by species in the surrounding electrolyte. In this work, we have specifically examined whether this occurs by detailed multiscale reaction–diffusion simulations that connect nanoscale events to macroscale observables that allow this specific reaction channel to be turned on and off. The mechanism is constructed using data from the literature, and provide a means of predicting system characteristics based on what is known. We find that recombination events are dominant in the photoanode, however they do not significantly affect photovoltage or photocurrent *operando* because the electrolyte chemistry has an adaptive response that maximizes the efficiency of the photoanode under all conditions. The resilience of this photoanode–electrolyte combination to losses may explain the success of the  $\text{I}^-/\text{I}_3^-$  redox couple for this class of solar energy conversion systems. This adaptive response works by coupling two  $\text{I}^-$  regeneration pathways that are accelerated together as the dye cycling rate and the associated electron loss rate increase. While previously identified adaptive chemistries involve responses to changes in structure, the chemistry identified here enables the system's function to adapt to changing conditions. The detailed predictions from the calculations provide a basis for testing them *via* experimental studies, which would be invaluable for refinement of our understanding of how inefficiencies in this important class of solar-driven processes arise.

## Author contributions

F. A. H. conceived of this study, performed all calculations and data analysis, and prepared the manuscript and figures.

## Funding source

This material is based upon work supported by the U.S. Department of Energy, Office of Science, Office of Basic Energy Sciences, Chemical Sciences, Geosciences, and Biosciences Division, in the Solar Photochemistry Program under Contract No. DE-AC02-05CH11231.

## Conflicts of interest

There are no conflicts of interest to declare.

## Acknowledgements

I thank Dr William D. Hinsberg (Columbia Hill Technical Consulting) for helpful discussions on simulation approaches during the course of this work, and Drs Thomas Cheshire and Peter Agbo (Lawrence Berkeley National Laboratory) and William D. Hinsberg for a critical reading of the manuscript. I

thank a Reviewer for suggesting bulk electrolyte spectroscopy studies as a test of simulation predictions.

## References

- 1 J. Bisquert, F. Fabregat-Santiago, I. Mora-Sero, G. Garcia-Belmonte and S. Gimenez, *J. Phys. Chem. C*, 2009, **113**, 17278–17290.
- 2 J. Bisquert, D. Cahen, G. Hodes, S. Ruhle and A. Zaban, *J. Phys. Chem. B*, 2004, **108**, 8106–8118.
- 3 J. Bisquert, A. Zaban, M. Greenshtein and I. Mora-Sero, *J. Am. Chem. Soc.*, 2004, **126**, 13550–13559.
- 4 S. Ardo and G. J. Meyer, *Chem. Soc. Rev.*, 2009, **38**, 115–164.
- 5 B. C. O'Regan, I. Lopez-Duarte, M. V. Martinez-Diaz, A. Forneli, J. Albero, A. Morandeira, E. Palomares, T. Torres and J. R. Durrant, *J. Am. Chem. Soc.*, 2008, **130**, 2906–2907.
- 6 F. Fabregat-Santiago, J. Bisquert, G. Garcia-Belmonte, G. Boschloo and A. Hagfeldt, *Sol. Energy Mater. Sol. Cells*, 2005, **87**, 117–131.
- 7 S. Y. Huang, G. Schlichthorl, A. J. Nozik, M. Gratzel and A. J. Frank, *J. Phys. Chem. B*, 1997, **101**, 2576–2582.
- 8 Y. C. Ling, J. K. Cooper, Y. Yang, G. M. Wang, L. Munoz, H. Y. Wang, J. Z. Zhang and Y. Li, *Nano Energy*, 2013, **2**, 1373–1382.
- 9 A. B. F. Martinson, T. W. Hamann, M. J. Pellin and J. T. Hupp, *Chem.–Eur. J.*, 2008, **14**, 4458–4467.
- 10 S. Ito, P. Liska, P. Comte, R. L. Charvet, P. Pechy, U. Bach, L. Schmidt-Mende, S. M. Zakeeruddin, A. Kay, M. K. Nazeeruddin and M. Gratzel, *Chem. Commun.*, 2005, 4351–4353, DOI: 10.1039/b505718c.
- 11 B. A. Gregg, F. Pichot, S. Ferrere and C. L. Fields, *J. Phys. Chem. B*, 2001, **105**, 1422–1429.
- 12 G. Boschloo and A. Hagfeldt, *Acc. Chem. Res.*, 2009, **42**, 1819–1826.
- 13 B. O'Regan, X. E. Li and T. Ghaddar, *Energy Environ. Sci.*, 2012, **5**, 7203–7215.
- 14 A. Hagfeldt, G. Boschloo, L. C. Sun, L. Kloo and H. Pettersson, *Chem. Rev.*, 2010, **110**, 6595–6663.
- 15 N. W. Duffy, L. M. Peter, R. M. G. Rajapakse and K. G. U. Wijayantha, *J. Phys. Chem. B*, 2000, **104**, 8916–8919.
- 16 J. Villanueva-Cab, H. X. Wang, G. Oskam and L. M. Peter, *J. Phys. Chem. Lett.*, 2010, **1**, 2800–2803.
- 17 S. A. Haque, Y. Tachibana, D. R. Klug and J. R. Durrant, *J. Phys. Chem. B*, 1998, **102**, 1745–1749.
- 18 A. Listorti, B. O'Regan and J. R. Durrant, *Chem. Mater.*, 2011, **23**, 3381–3399.
- 19 Y. Tachibana, S. A. Haque, I. P. Mercer, J. R. Durrant and D. R. Klug, *J. Phys. Chem. B*, 2000, **104**, 1198–1205.
- 20 Y. Tachibana, S. A. Haque, I. P. Mercer, J. E. Moser, D. R. Klug and J. R. Durrant, *J. Phys. Chem. B*, 2001, **105**, 7424–7431.
- 21 H. X. Wang and L. M. Peter, *J. Phys. Chem. C*, 2012, **116**, 10468–10475.
- 22 J. Bisquert and I. Mora-Sero, *J. Phys. Chem. Lett.*, 2010, **1**, 450–456.
- 23 J. J. Lee, G. M. Coia and N. S. Lewis, *J. Phys. Chem. B*, 2004, **108**, 5269–5281.



- 24 A. N. M. Green, R. E. Chandler, S. A. Haque, J. Nelson and J. R. Durrant, *J. Phys. Chem. B*, 2005, **109**, 142–150.
- 25 T. J. Barr and G. J. Meyer, *ACS Energy Lett.*, 2017, **2**, 2335–2340.
- 26 C. E. Richards, A. Y. Anderson, S. Martiniani, C. H. Law and B. C. O'Regan, *J. Phys. Chem. Lett.*, 2012, **3**, 1980–1984.
- 27 P. R. F. Barnes, A. Y. Anderson, J. R. Durrant and B. C. O'Regan, *Phys. Chem. Chem. Phys.*, 2011, **13**, 5798–5816.
- 28 F. A. Houle, *J. Phys. Chem. C*, 2019, **123**, 14459–14467.
- 29 S. Macnamara, K. Burrage and R. B. Sidje, *Multiscale Model. Simul.*, 2008, **6**, 1146–1168.
- 30 W. D. Hinsberg and F. A. Houle, *Kinetiscope*, available at <http://www.hinsberg.net/kinetiscope>, accessed May 12, 2019.
- 31 D. L. Bunker, B. Garrett, T. Kleindienst and G. S. Long, *Combust. Flame*, 1974, **23**, 373–379.
- 32 D. T. Gillespie, *J. Comput. Phys.*, 1976, **22**, 403–434.
- 33 W. D. Hinsberg and F. A. Houle, Software Architecture for Stochastic Simulation of Non-Homogeneous Systems, *US Pat.*, 5826065, 1998.
- 34 F. A. Houle, W. D. Hinsberg, M. Morrison, M. I. Sanchez, G. Wallraff, C. Larson and J. Hoffnagle, *J. Vac. Sci. Technol., B: Microelectron. Nanometer Struct.–Process., Meas., Phenom.*, 2000, **18**, 1874–1885.
- 35 T. P. Cheshire, M. K. Brennaman, P. G. Giokas, D. F. Zigler, A. M. Moran, J. M. Papanikolas, G. J. Meyer, T. J. Meyer and F. A. Houle, *J. Phys. Chem. B*, 2020, **124**, 5971–5985.
- 36 F. A. Houle, A. A. Wiegel and K. R. Wilson, *J. Phys. Chem. Lett.*, 2018, **9**, 1053–1057.
- 37 A. Hauch and A. Georg, *Electrochim. Acta*, 2001, **46**, 3457–3466.
- 38 L. Bay, K. West, B. Winther-Jensen and T. Jacobsen, *Sol. Energy Mater. Sol. Cells*, 2006, **90**, 341–351.
- 39 J. G. Rowley, B. H. Farnum, S. Ardo and G. J. Meyer, *J. Phys. Chem. Lett.*, 2010, **1**, 3132–3140.
- 40 F. G. K. Baucke, R. Bertram and K. Cruse, *J. Electroanal. Chem.*, 1971, **32**, 247–256.
- 41 A. J. Parker, *J. Chem. Soc. A*, 1966, 220–228.
- 42 M. F. Ruasse, J. Aubard, B. Galland and A. Adenier, *J. Phys. Chem.*, 1986, **90**, 4382–4388.
- 43 J. van de Lagemaat and A. J. Frank, *J. Phys. Chem. B*, 2000, **104**, 4292–4294.
- 44 B. C. O'Regan and J. R. Durrant, *Acc. Chem. Res.*, 2009, **42**, 1799–1808.
- 45 J. Nissfolk, K. Fredin, A. Hagfeldt and G. Boschloo, *J. Phys. Chem. B*, 2006, **110**, 17715–17718.
- 46 L. M. Peter, *J. Phys. Chem. C*, 2007, **111**, 6601–6612.
- 47 A. V. Barzykin and M. Tachiya, *J. Phys. Chem. B*, 2004, **108**, 8385–8389.
- 48 T. W. Hamann, R. A. Jensen, A. B. F. Martinson, H. Van Ryswyk and J. T. Hupp, *Energy Environ. Sci.*, 2008, **1**, 66–78.

