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Polymorphism in metal halide perovskites

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Metal halide perovskites (MHPs) are frontrunners among solution-processable materials for lightweight, large-area and flexible optoelectronics. These materials, with the general chemical formula AMX3, are structurally complex, undergoing multiple polymorph transitions as a function of temperature and pressure. In this review, we provide a detailed overview of polymorphism in three-dimensional MHPs as a function of composition, with $A = Cs^+$, MA^+ , or FA^+ , $M = Pb^{2+}$ or Sn^{2+} , and $X = Cl^-$, Br^- , or I^- . In general, perovskites adopt a highly symmetric cubic structure at elevated temperatures. With decreasing temperatures, the corner-sharing MX₆ octahedra tilt with respect to one another, resulting in multiple polymorph transitions to lower-symmetry tetragonal and orthorhombic structures. The temperatures at which these phase transitions occur can be tuned via different strategies, including crystal size reduction, confinement in scaffolds and (de-)pressurization. As discussed in the final section of this review, these solid-state phase transformations can significantly affect optoelectronic properties. Understanding factors governing these transitions is thus critical to the development of highperformance, stable devices.

Introduction

As conversion efficiencies of metal halide perovskite (MHP) solar cells skyrocket past 25%, research in the fundamental parameters governing MHP performance in optoelectronic devices, including solar cells, light-emitting diodes, transistors, and sensors, continues to rise exponentially. This class of materials, adopting the general perovskite composition of AMX₃ where X represents halide ions of Cl⁻, Br⁻ or I⁻, exhibit superior light absorption and charge mobilities compared to their organic molecular and polymeric counterparts, 2 as well as remarkable defect tolerance that enables their high performance upon solution processing.^{3,4}

Intricately tied to the performance of these materials is the specific arrangement adopted by the ions within the crystalline lattice. All MHPs exhibit different polymorphs, i.e. solid-state crystal structures, as a function of temperature and pressure. The optoelectronic properties of MHPs can vary significantly depending on the specific polymorph present. CsPbI₃, FAPbI₃ and CsSnI₃, for example, all adopt a non-perovskite, insulating phase at room temperature that cannot be used as active layers in optoelectronic devices.^{5–8} Significant research effort has thus focused on developing methods to trap the metastable, active perovskite phases of these compounds.^{9,10} In other systems, the relationship between polymorphism and optoelectronic properties is less drastic, but important nonetheless. The MAPbI₃

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tetragonal-to-cubic transition, for example, occurs within the operating range of solar cells around 330 K and can impact, for example, charge carrier dynamics. 11 Understanding and controlling polymorphism in this class of materials is necessary to ensure long-lasting, reliable performance of MHP-based optoelectronic devices.

In this review, we aim to present a systematic and comprehensive picture of polymorphism in three-dimensional MHPs, beginning with a general classification of perovskite polymorphs and a comparison of all the known polymorph transition temperatures across 17 different three-dimensional MHPs. These MHPs represent "pure" compositions, rather than mixed compositions in which one or all of the components is a mixture of different ionic species (e.g. MA_rCs_{1-r}-Pb_νSn_{1-ν}Cl_zBr_{3-z}). Overall, all MHPs adopt a high-symmetry cubic phase at elevated temperatures, transitioning to lowersymmetry polymorphs upon decreasing temperature as a result of octahedral tilting. For MHPs with organic A-site cations, the orientation and motion of the cation with respect to the metalhalide framework must also be considered. Strategies, including crystal size reduction and pressurization, employed to shift polymorph transition temperatures and relative thermodynamic stabilities are discussed. An overview of how optoelectronic properties vary with different polymorphs is also provided.

Classifying perovskite polymorphs

First discovered by Russian mineralogist Lev Perovski, perovskites are a class of materials that adopt the structure of calcium titanium oxide, with a general chemical formula of AMX3.

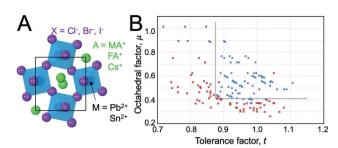


Fig. 1 (A) Perovskite structure comprising corner-sharing octahedra. (B) Plot of octahedral factors versus tolerance factors based on anion-dependent cationic radii. Blue circles represent compounds that form the perovskite structure at room temperature and pressure, while red x's represent those that do not. Adapted from ref. 15. Copyright 2016, Royal Society of Chemistry.

In the perovskite structure, X ions form corner-sharing octahedra with M ion centers, while A ions occupy the spaces between octahedra. A schematic of the perovskite structure highlighting the corner sharing octahedra is provided in Fig. 1A. This particular structure represents an orthorhombic phase in which the octahedra are tilted relative to one another. The extent and direction of octahedral tilting define the different perovskite polymorphs and are a consequence of a mismatch between the sizes of the ions comprising the perovskite structure. 12 The ideal size of the A-site cation can be predicted by the Goldschmidt tolerance factor, t:13

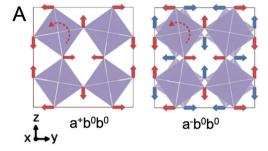
$$t = \frac{r_{\rm A} + r_{\rm X}}{\sqrt{2}(r_{\rm M} + r_{\rm X})}\tag{1}$$

where r_A , r_M , and r_X , represent the radii of the A, M and X ions, respectively. For MHPs, the size of the A cation is generally smaller than required for t = 1.¹⁴ Recently, the use of anion-dependent cation radii to more accurately predict perovskite formation has been proposed¹⁵ in conjunction with the octahedral factor, μ , defined as:16

$$\mu = \frac{r_{\rm M}}{r_{\rm X}} \tag{2}$$

As displayed in Fig. 1B, Travis and co-workers found that combining revised tolerance factors and octahedral factors led to the accurate prediction of whether or not compounds form perovskite structures at room temperature and pressure based on threshold values of $t \le 1.06$ and $\mu > 0.41.^{15}$

At ambient pressure, perovskites can undergo multiple solid-state phase transitions as a function of temperature. In the highest-symmetry cubic structure with a M-X-M angle of 180° that appears at elevated temperatures, the corner-sharing octahedra are aligned in all directions. With decreasing temperature, the octahedra tilt due to a mismatch between the size of the A-site cation and the metal-halide framework. Polymorphs thus represent different extents of tilting of the corner-sharing octahedra. The tilting can be described in Glazer notation, ¹⁷ which assigns either an in-phase (+) or outof-phase (-) rotation of octahedra along each of the three spatial dimensions, x (first position), y (second position), and z (third position). The amplitude of these rotations is indicated



B

a

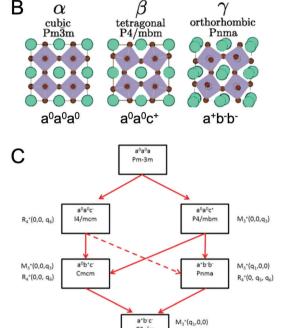


Fig. 2 (A) Illustrations of different octahedral tilting directions and their corresponding identifier in Glazer notation. (B) Solid-state phases observed in MHPs, from the high-temperature, high-symmetry cubic structure to the low-temperature, low-symmetry orthorhombic structure. Adapted with permission from ref. 12. Copyright 2018, American Physical Society. (C) Flow diagram of phase transitions in organic MHPs and all-inorganic MHPs. The dashed red line indicates that transitions from 14/mcm to Pnma symmetry must be a first order transition. Reproduced with permission from ref. 18. Copyright 2016, Royal Society of Chemistry.

R4+(0, q5, q6)

by lower case letters. As detailed by Bechtel and co-workers in Fig. 2A, a perovskite structure characterized as a⁺b⁰b⁰ in Glazer notation exhibits in-phase octahedral tilting along the x-axis, with no tilting along other axes.12 On the other hand, a structure characterized as a b b exhibits out-of-phase octahedral tilting in the x direction. For a structure characterized by a b, in-plane octahedral rotation occurs along the x-direction, while out-of-plane tilting occurs along the y- and z-directions. The amplitude of tilting is the same in the γ - and z-directions, but different from the x-direction, as indicated by the different lowercase letters a and b. In general, perovskites undergo a cubic-to-tetragonal transition, followed by a tetragonal-to-orthorhombic transition with decreasing temperature.

The cubic, tetragonal, and orthorhombic phases are represented by the Greek letters α , β and γ , respectively. Fig. 2B displays the general classification for these phases. 18 For all MHPs characterized thus far, octahedral tilting in the γ -phase is characterized by a b, resulting in *Pnma* symmetry, while the tetragonal phase can adopt either P4/mbm symmetry with in-phase octahedral tilting denoted by a⁰a⁰c⁺ or I4/mcm symmetry with out-ofphase octahedral tilting denoted by $a^0a^0c^-$. Fig. 2C displays a flow diagram of phase transitions observed in hybrid and allinorganic perovskites, including those outside of the halide perovskite family.18

For the sub-class of optoelectronically-active perovskites that are the topic of this review, the use of halide anions (e.g. Cl⁻, Br⁻, I⁻) in the X position imposes restrictions on the options for A and M cations. Compared to other anions common to perovskite structures, such as O2-, halide anions have smaller negative charge and larger ionic radii. 19 As such, A-site cations have thus far been limited to organic methylammonium and formamidinium and inorganic cesium, while MII cations are largely restricted to heavier group IV elements (e.g. Sn²⁺ and Pb²⁺).¹⁹

Fig. 3 displays the solid-state phase transition temperatures and space groups reported for bulk MHP crystals, including $\begin{array}{l} \text{CsPbCl}_{3},^{20-23} \text{ CsPbBr}_{3},^{24-26} \text{ CsPbI}_{3},^{27,28} \text{ CsSnCl}_{3},^{29,30} \text{ CsSnBr}_{3},^{31,32} \\ \text{CsSnI}_{3},^{29,33,34} \text{ MAPbCl}_{3},^{35-37} \text{ MAPbBr}_{3},^{36,38-40} \text{ MAPbI}_{3},^{36,41-43} \end{array}$ MASnCl₃,³⁰ MASnBr₃,⁸ MASnI₃,^{44,45} FAPbCl₃,⁴⁶ FAPbBr₃,⁴⁷ FAPbI₃,^{6,8} FASnBr₃,⁴⁸ and FASnI₃.⁴⁷ To the best of our knowledge, the crystal structure of FASnCl3 has not been reported. As observed from the figure, all perovskites adopt a cubic $Pm\bar{3}m$ structure at elevated temperatures. In general, chlorine-based derivatives retain the cubic structure to relatively lower temperatures compared to bromine and iodine derivatives with the same A-site and metal cations. With a smaller ionic radius, the tolerance factors for chlorine derivatives are closer to 1, rendering the cubic phase more stable. For all derivatives, the symmetry of the crystal structure decreases as temperature is lowered, following the flow diagram displayed in Fig. 2C. Furthermore, MHPs with organic A-site cations generally form the cubic phase at room temperature due to the relatively larger size of these cations compared to Cs⁺. For reference, Cs⁺ has an ionic radius of 1.88 Å. 15 Estimated effective ionic radii generally fall in the range of 2.03-2.38 Å for MA⁺ and 2.24-2.53 Å for FA⁺. 15,49-51

We note there are some discrepancies in the literature on phase transition temperatures and space groups observed for specific compositions. In 1987, for example, Poglitsch and Weber reported the space groups of the low temperature phases of MAPbCl₃, MAPbBr₃ and MAPbI₃ to be P222₁, Pna2₁, and Pna21, respectively. 36 These assignments have been reassigned to Pnma based on recent high resolution X-ray and neutron diffraction experiments. 35,43,52 For MAPbBr₃, there are further conflicting reports about the existence of a P4/mmm tetragonal phase in between the orthorhombic phase and I4/mcm tetragonal phase, with neutron scattering⁵² and X-ray diffraction experiments⁵³ unable to detect this phase.

Polymorphism in all-inorganic perovskites

For perovskites comprising halide anions in the X-site, cesium is the only elemental ion with a sufficiently large radius to support the perovskite framework. 19 To date, the most promising inorganic halide perovskites for optoelectronic devices are CsPbX₃ and CsSnX₃ where X = Cl⁻, Br⁻, or I⁻. The first reports on the crystal structures of these compounds date back to the late 1950s to early 1970s, 29,55-57 but it was not until 2015 that the first all-inorganic light-emitting diodes comprising CsPbBr₃ as the photoactive layer⁵⁸ and solar cells comprising CsPbI₃ as the photoactive layer⁵⁹ were successfully fabricated. Since then, inorganic MHP-based devices have been demonstrated for a variety of devices, 60,61 including photodetectors, 62,63 solar cells, 64,65 transistors 66,67 and light-emitting diodes. 68-70

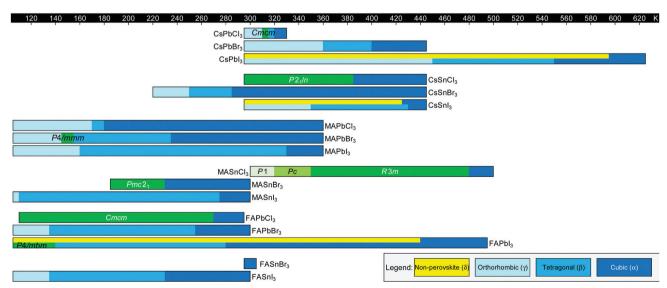


Fig. 3 Summary of phase transition temperatures in halide perovskites reported in the literature collected at ambient pressure as a function of composition. $6.8.20^{-48}$ Cubic structures exhibit $Pm\bar{3}m$ symmetry, tetragonal structures exhibit either I4/mcm or P4/mbm symmetry, and the orthorhombic phase exhibits Pnma symmetry, unless otherwise noted. Green colors indicate phases not represented by the above categories.

Promisingly, CsPbI3-based solar cells recently achieved record efficiencies of 19%.71

Compared to MHPs utilizing organic molecules as the A-site cation, inorganic MHPs comprising inorganic Cs⁺ cations are expected to exhibit enhanced thermal stability.⁵⁴ However, these inorganic perovskites have other critical issues related to degradation and polymorph transitions. Sn-Based perovskites exhibit smaller bandgaps compared to their Pb-based counterparts, 1.3 eV for CsSnI₃ compared to 1.7 eV for CsPbI₃, but suffer from rapid oxidation of Sn²⁺ to the more stable state of Sn⁴⁺. In the Pb family, CsPbCl₃ and CsPbBr₃ are air-stable in the perovskite phase at ambient temperatures and pressures but their relatively large bandgaps of 2.99 and 2.31 eV⁷² make them less attractive for light harvesting applications. On the other hand, the perovskite phases of CsPbI₃ possess reasonable bandgaps around 1.7 eV but these phases are only stable at elevated temperatures. Of all the CsMX₃ compounds, CsPbI₃ and CsSnI₃ have the smallest tolerance factors of 0.89 and 0.91 (based on anion-dependent A-site cation radii) and octahedral factors of 0.47 and 0.44, respectively. 15 This mismatch between the A-site cation size and the perovskite framework renders these systems less stable compared to other CsMX₃ compounds where $M = Pb^{2+}$ or Sn^{2+} and $X = Cl^{-}$ or Br^{-} . At room temperature, CsPbI3 and CsSnI3 both exist in a yellow, non-perovskite polymorph, 55,73 referred to as the δ phase, exhibiting large bandgaps of around 2.8 and 2.55 eV, respectively. 34,74 The δ phase is characterized by isolated double MX₆ octahedra (Fig. 4A).

While all four phases of CsSnI₃ were identified and characterized^{29,75} decades prior to the fabrication of the first perovskite-based solar cell by Kojima and coworkers in 2009,⁷⁶ some confusion initially surrounded the room temperature perovskite structure of CsPbI3. Many early reports on CsPbI3based optoelectronic devices described the room temperature metastable perovskite phase as the cubic α phase. Recent studies, however, have found that CsPbI3 undergoes the same cubic-tetragonal and tetragonal-orthorhombic phase transitions upon cooling as other perovskites.^{27,28} Fig. 4A displays the phase transitions of CsPbI3 as a function of the thermal history. At room temperature, CsPbI₃ adopts the non-perovskite δ phase. Upon heating above 587 K, this phase converts directly to the cubic α phase.²⁷ If the α phase is then slowly cooled, it undergoes a transition to the tetragonal β phase at 554 K and then to the orthorhombic y phase at 457 K. Comparing two different methods of stabilizing the black phase of CsPbI₃ at room temperature via Rietveld refinement of powder diffraction patterns, Sutton and coworkers confirmed that the γ phase is present at room temperature, not the cubic α phase.²⁸

At room temperature, the γ phases of both CsSnI₃ and CsPbI3 are unstable and will spontaneously convert to the insulating δ phases,⁷⁵ even when stored under vacuum.⁷⁷ Using low-frequency Raman spectroscopy to monitor phonon modes in a CsPbI3 crystalline film deposited via drop casting from dimethyl sulfoxide onto a SiO2 substrate, we found this transition to occur within 6 min of air exposure (Fig. 4B).⁷⁸ The transition itself was rapid and completed within time resolution limit of the measurements (e.g. within one minute).

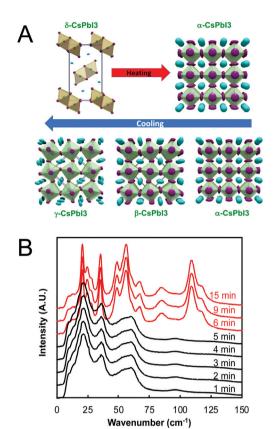


Fig. 4 (A) Temperature-dependent phase transitions of CsPbl₃ at ambient pressure. Reproduced with permission from ref. 27. Copyright 2018, American Chemical Society. (B) Room-temperature low-frequency Raman spectra of a drop cast CsPbl₃ film as a function of exposure time to air $(\lambda_{ex} = 976 \text{ nm})$. CsPbI₃ initially adopted the γ phase (black) upon cooling from 610 K before converting to the δ phase (red) between 5 and 6 min of air exposure. Adapted with permission from ref. 78. Copyright 2020, Royal Society of Chemistry.

Stabilizing the active black phases of these compounds at room temperature has thus been an area of intense research focus over the past five years.

Polymorphism in organometallic perovskites

Organic MHPs with methylammonium (MA+) and formamidinium (FA+) as the A-site cation represent the most extensively studied compositions of MHPs for optoelectronic applications to-date. These ions are larger than Cs⁺, resulting in tolerance factors closer to, and in some cases larger than 1.79 As displayed in Fig. 3, phase transitions in organic MHPs are similar to inorganic perovskites. These compounds adopt a high symmetry cubic phase at elevated temperatures and transition to lower symmetry structures via octahedral tilting upon cooling. Unlike inorganic MHPs with a simple atomic A-site cation, however, phase transitions in organic MHPs possess additional complexity associated with orientation and rotation of the asymmetric molecular cations in the interstices of the cornersharing MX₆ octahedra. 80 The specific orientations adopted by the A-site cation may in turn affect optoelectronic properties, such as the bandgap.81

Methylammonium lead iodide (MAPbI₃). MAPbI₃, one of the intensely studied compositions among halide perovskites, was first reported by Weber in 1947.82 This compound undergoes two phase transitions upon cooling, a cubic-to-tetragonal transition at 330 K and a tetragonal to orthorhombic transition at 160 K.83 In each of these phases, the MA+ cation is surrounded by a different environment. The MA⁺ cation itself is asymmetric and can rotate about and along the C-N axis. In recent years, rotational dynamics of the MA⁺ cation in these phases have been studied by calorimetry, 84,85 dielectric spectroscopy, 85-87 computational methods, 88-92 quasielastic and inelastic neutron scattering^{93–97} nuclear magnetic resonance (NMR) spectroscopy 98-101 infrared spectroscopy, 83,84,102 isotopic substitution, 103 and Raman spectroscopy. 104

In the cubic phase, which is stable above 330 K, early magnetic resonance experiments found that the MA+ cation can rotate on the picosecond timescale within the interstitial spaces between corner-sharing PbI₆ octahedra. In 2015, Weller and co-workers used neutron powder diffraction to provide a comprehensive picture of MA+ cation dynamics in the range of 100-352 K. 43 Fig. 5A displays a MAPbI₃ unit cell in the cubic phase, with atomic displacement parameter (ADP) ellipsoids representing a 50% probability of locations of the hydrogen atoms at the ends of the MA⁺ cations. As observed from the figure, the MA⁺ ion appears disordered in the cubic phase. Later studies revealed preferential alignment of the MA⁺ cations based on four-fold rotational symmetry of the C-N axis and three-fold rotational symmetry about the C-N axis. 97,106,107 Employing single crystal neutron diffraction studies, Ren and coworkers found that the MA⁺ cations preferentially align along the [011] (edge), [111] (diagonal), and [100] (face) directions. 108 These experiments also revealed that the middle point of the C-N bond is located slightly off-center due to interactions between the amine group of MA⁺ and iodine atoms.

Upon decreasing the temperature below 330 K, MAPbI₃ undergoes a phase transition to the tetragonal β phase. Here the PbI₆ octahedra are tilted compared to the cubic phase, with an average Pb-I-Pb bond angle of 165.3°. Fig. 5b displays one configuration of the tetragonal phase at 180 K for MAPbI3 with ADP ellipsoids at 50% probability. 43 In this phase, the NH₃ group adopts four distinct positions, while the CH₃ group adopts eight distinct positions. 108 Although the organic cation is more confined in the tetragonal phase compared to cubic phase, and exhibits slower rotational dynamics, 105 four-fold rotation of (C-N) axis C₄ and three-fold rotation around (C-N) axis C₃ is still possible. 97,106,107 Elastic and quasi-elastic neutron scattering techniques and group theoretical analysis revealed the room temperature relaxation times of and around the C-N axis to be approximately 5 and 1 ps, respectively. 94,106

It should be noted that the crystal structure classification of tetragonal MAPbI3 is still being debated, in part due to extensive twinning that complicates structure determination. 109,110 A centrosymmetric I4/mcm space group for tetragonal MAPbI₃ was proposed by Poglitsch and Weber in 1987.³⁶ In 2013, Stoumpos and coworkers reported instead a non-centrosymmetric I4cm space group due to off-centering of the MA+ cation along the c-axis. 111 A later report suggested that noncentrosymmetric tetragonal MAPbI3 was achieved through static distortion of the Pb-I framework via second harmonic generation analysis. 112 Subsequent second harmonic generation rotational anisotropy experiments capable of distinguishing surface and bulk signals, however, determined that tetragonal MAPbI₃ indeed adopts the centrosymmetric I4/mcm space group. 113 Structural rearrangement at perovskite surfaces may lead to the formation of noncentrosymmetric domains that are not present in the bulk, as observed in low-temperature scanning tunneling microscopy images of MAPbBr₃ crystals. 114 Noncentrosymmetric structures in MHPs are expected to exhibit ferroelectricity and Rashba splitting due to spin-orbit coupling interactions with charge carriers. With significant implications to fundamental device physics and applications, intensive research efforts continue to focus on definitive structure determination in these materials.

As displayed in Fig. 5C, the MA⁺ cation is in its most confined state in the low temperature orthorhombic γ phase (T < 160 K), for which the average Pb-I-Pb bond angle of 154.51°. In this phase, the MA+ cations are completely ordered in an alternating head-to-tail configuration with the NH3 groups of MA⁺ closer to the inorganic framework compared to the CH₃ groups. 43 Relaxation times of the MA+ cation were found to be orders of magnitude larger than those expected for a rigid rotator in the same temperature range due to restrictions on motion

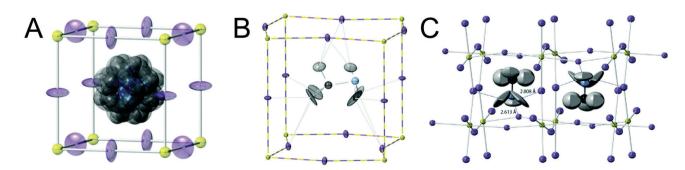


Fig. 5 Depictions of MA⁺ cation orientations in the (A) cubic phase with ADP ellipsoids at 50% probability, (B) one orientation of the tetragonal phase with ADP ellipsoids at 50% probability, and (C) orthorhombic phase with ADP ellipsoids at 90% probability of MAPbI₃, as determined by powder neutron diffraction experiments. Adapted with permission from ref. 43. Copyright 2015, Royal Society of Chemistry.

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imposed by the Pb-I framework. 115 This configuration promotes hydrogen bonding between the NH₃ group of the MA⁺ cation and the framework. 43,116 While four-fold rotation is not allowed in this polymorph, three-fold rotation about the C-N axis is possible, 97,106,107 with estimated activation energies of 56 meV and 120 meV for rotations of the CH3 and NH3 tails, respectively. 117 Using infrared spectroscopy to examine vibrational modes in MAPbI₂ single crystals. Schuck and coworkers found that the vibrational mode associated with symmetrical NH₃ bending shifted to lower frequencies upon transitioning from the tetragonal to orthorhombic structure due to hydrogen bonding between the amine group on the organic cation and halide ions. 118 Interestingly, the strength of the interaction depended on the halide composition, with the largest effect observed for I compared to Cl and Br.

These experimental results have been supported by extenmodelling of MA+ computational dynamics. 18,94,95,102,119-122 Recent molecular dynamics simulations of phase transitions in organic MHPs have found that phase transitions are the result of dipole-dipole interactions between organic cations and hydrogen bonding with the Pb-I lattice. 116 Reduced cation movement with decreasing temperature increases the strength of these interactions, resulting in lattice deformation. 121 Combining DFT calculations and ab initio molecular dynamics, Deretzis and Magna found that vibrational features of MAPbI3 are strongly affected by the orientation of the organic cation in the low temperature orthorhombic phase, with the tetragonal-orthorhombic phase transition occurring to relax stress via rearrangement of the MA⁺ cation in an orientation that can reduce local stretching of ionic bonds in the inorganic framework. 120

Formamidinium lead iodide. The second most widely studied organic cation used in three-dimensional MHPs is formamidinium (FA⁺).^{123,124} With a bulkier size compared to MA+, FA+-based MHPs exhibit a smaller bandgap and improved thermal and lattice stability. 125,126 In general, FA-based MHPs undergo similar temperature-dependent crystal structure transitions as MA-based MHPs, 127 with the major exception being that FAPbI₃ typically exists in a non-perovskite δ phase at room temperature with P63mc symmetry. 42 Calculated ground state energy differences between δ and α -FAPbI₃ phases have demonstrated that the δ phase is the thermodynamically-stable phase due to a face-sharing configuration of PbI₆ octahedra that results in higher structural stability compared to cornersharing octahedra in the α phase. Because the α - δ phase transition in FAPbI₃ is reversible in solution, however, the α phase can be kinetically trapped under appropriate processing conditions⁴² to be incorporated into optoelectronic devices.

For FA⁺ cations, there are three primary axes of rotation, with rotation about the N-N axis the dominant motion in all phases above 50 K. 107 As displayed in Fig. 6A, FA cations are completely disordered in the high temperature cubic phase and can rotate freely within the inorganic framework. 127 In this phase, the carbon atom of the FA⁺ cation is located at the center of the inorganic framework. 128 In the lower temperature tetragonal phase, FA+ cations rotate between preferred orientations,

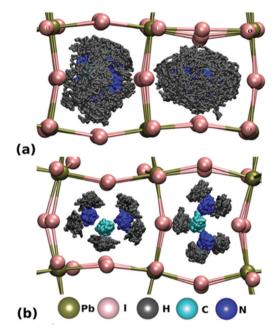


Fig. 6 Molecular dynamics simulations of FA cations within Pb-I frameworks in the (A) pseudocubic phase and (B) tetragonal phase. Reproduced with permission from ref. 127. Copyright 2018, American Chemical Society.

as displayed in Fig. 6B. In this phase, FA⁺ cations in adjacent inorganic cages reorient with respect to each other by changing the H-C bond direction. This preferred orientation is influenced by PbI6 octahedral tilting, as demonstrated by both experimental and simulation results. 127,129 In the low-temperature orthorhombic phase, FA⁺ cation rotation is arrested but unlike MA⁺ cations, FA⁺ cations do not exhibit any long range order.127 Instead, their configuration can be described by a glass-like, disordered state. 107

Influence of crystal size on polymorph transitions

One of the most promising emerging strategies to stabilize MHPs is to reduce the average crystal size. 130 The total Gibbs free energy of a crystal, $G_{\rm T}$, can be expressed as a sum of the volume free energy, $G_{\rm V}$, and the total surface free energy: 131

$$G_{\rm T} = G_{\rm V} + \left(\frac{a}{\rho}\right) \frac{A}{V} \tag{3}$$

where a is the surface energy of the crystal surface, ρ is the bulk crystal density, A is the mean total surface area and V is the mean volume per crystal. As crystals grow the surface area-tovolume ratio, A/V, decreases. For bulk crystals, the surface free energy contribution is thus minimal. For bulk crystals, the Gibbs free energy is largely determined by the volume free energy.

For relatively small crystals, on the other hand, the surface free energy contribution to the total Gibbs free energy can be significant. Fig. 7 illustrates the shift in Gibbs free energy upon crystal size reduction of two polymorphs of a compound as a function of temperature. As can be seen from the figure, at temperatures above $T_{\alpha\beta}$, the α phase is favored due to its lower Gibbs free energy compared to the β phase. $T_{\alpha\beta}$ represents the

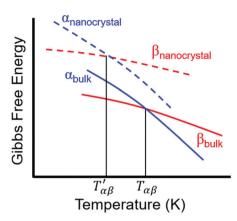


Fig. 7 Comparison of the dependence of the Gibbs free energy on temperature for different polymorphs of a bulk crystal and nanocrystal. Adapted with permission from ref. 130. Copyright 2019, American Chemical Society.

temperature at which phase co-existence occurs. Below $T_{\alpha\beta}$, the β phase is thermodynamically preferred. For nanocrystals with large A/V ratios (represented by dashed lines), the Gibbs free energy of both phases increases due to an increased contribution from the surface free energy. If the surface free energy of the β phase is larger than the surface free energy of the α phase, the final result is a lowering of the phase transition temperature from $T_{\alpha\beta}$ to $T'_{\alpha\beta}$. Indeed, lower surface free energies of higher symmetry polymorphs compared to lower symmetry polymorphs have been recorded for a variety of metal oxide systems. 132-135 For CsPbI₃, DFT calculations found the surface energies of the high symmetry γ phase and low symmetry δ phase to be 0.13 and 2.57 J m⁻², respectively. 136

Perovskite quantum dots (PQDs)

At the smallest length scale, many research efforts have focused on the synthesis of perovskite quantum dots (PQDs) with dimensions in the range of 2-20 nm. 137,138 Lowering of solidstate phase transition temperatures in these systems has been widely reported, although some confusion has surrounded the precise solid-state phases adopted by PQDs. In the first report of metal halide PQDs in 2015, Protesescu and coworkers synthesized all-inorganic cesium lead halide perovskite QDs (CsPbX₃, X= Cl, Br, I, and mixed Cl/Br, Br/I), with sizes in the range of 4-15 nm. XRD patterns collected on the PQDs were used to assign the cubic α phase to all compositions. These phases were reported to be stable, with only CsPbI₃ transitioning to the orthorhombic, non-perovskite δ phase upon several months of storage. A later report using synchrotron radiation and atomic pair distribution (PDF) to characterize both elastic and inelastic X-ray scattering from PQDs prepared in the same manner, however, found that CsPbCl₃, CsPbBr₃ and CsPbI₃ QDs all existed in the room temperature γ phase observed in bulk crystals. ¹⁴⁰ Here the authors found that the presence of twin boundaries in some cases gave rise to "apparent higher symmetry" observed via X-ray diffraction. In general, while the main peaks in the XRD patterns of the γ phase are similar to those of the α phase,

the presence of weak superstructure peaks indicate octahedral tilting in the perovskite framework. These findings are supported by ab initio DFT simulations that found crystal size reduction to below 5.6 nm and 2.7 nm were needed to thermodynamically stabilize γ-CsPbI₃ and α-CsPbI₃, respectively. 141

In 2018, Cottingham and Brutchey experimentally examined phase transitions in CsPbBr3 PQDs with average diameters of 9 nm using temperature-dependent, synchrotron radiation-based total X-ray scattering techniques. 142 At room temperature, these PQDs were found to exist in the γ phase. ¹⁴³ The γ - β transition was observed between 50–59 °C, while the β - α transition was observed between 108-117 °C. These ranges represent phase transition temperature depressions of at least 29 and 13 °C, respectively, compared to those observed in bulk crystals. In the analogous organic-inorganic hybrid system of MAPbBr3, Liu and coworkers used temperature-dependent photoluminescence measurements to identify the β - α transition for PQDs with an average diameter of 3.7 ± 0.6 nm at 150 K, ¹⁴⁴ compared to \sim 236 K in the bulk. ⁴⁰ Close agreement was found between the experimental data and the predicted phase transition temperature between the paraelectric α phase and ferroelectric β-phase following Ginzburg-Landau-Devonshire theory. Fig. 8 displays the size dependence of the critical temperature and Gibbs free energy, decrease and increase, respectively, with decreasing crystal size. As observed from the figure, the critical temperature drops to 0 K for particle diameters below 2 nm, in agreement with the authors' observation of the absence of phase transitions for PQDs with average diameters of $1.7 \pm 0.4 \text{ nm}.$

Coexistence of the γ and α phases of CsPbBr₃ PQDs has also been reported. Recently, Brennan and coworkers examined the crystal structure of CsPbBr3 nanocubes with average edge lengths 5 nm and 10 nm via lattice-resolved transmission electron microscopy imaging at progressive defocus values. 145 Finite Fourier transforms of experimental images and simulated images of the α and γ phases suggest that 10 nm diameter CsPbBr₃ nanocubes mainly adopt the cubic α phase, with some γ phase present. For 5 nm diameter nanocubes, on the other hand, only the α phase was found to be present.

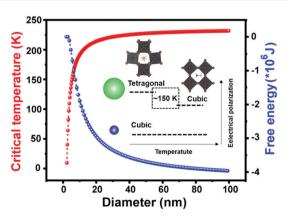


Fig. 8 Critical temperature and Gibbs free energy versus particle diameter for MAPbBr₃ PQDs based on Ginzburg-Landau-Devonshire theory. Reproduced with permission from ref. 144. Copyright 2019, American Chemical Society.

Scaffold-confined crystals

Review

Another method of controlling polymorphism via crystal size reduction is to confine crystallization within nanoporous scaffolds, such as controlled porous glass, selectively-etched block copolymers and anodized aluminium oxide. 130,146 This strategy has been successfully employed to shift phase transition temperatures, including melting points, in a number of different organic and inorganic systems, as summarized in several reviews. 130,147,148

Recently, our group examined the phase stability of MAPbI₃ crystals confined within the cylindrical nanopores of anodized aluminium oxide scaffolds with average pore diameters of 100-250 nm (Fig. 9A). 149,150 Temperature-dependent X-ray diffraction measurements in the range of 293-373 K and photoluminescence measurements in the range of 4-300 K revealed the α - β transition temperature to shift from 330 K for bulk crystals to 170 K for nanoconfined crystals. Similarly, the β - γ transition was observed at 80 K for nanoconfined crystals, compared to 150 K in the bulk. Fig. 9B displays a comparison of the Gibbs free energy curves as a function of

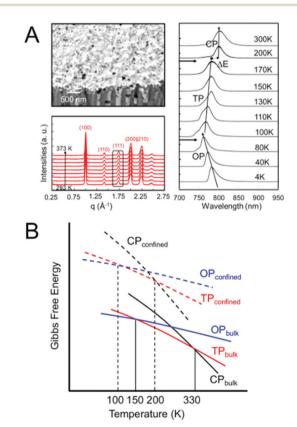


Fig. 9 (A) SEM image of MAPbI₃ confined within 100 nm diameter nanopores of anodized aluminum oxide templates. Temperaturedependent X-ray diffraction measurements revealed only the cubic phase to be present in the range of 293-373 K, while temperature-dependent photoluminescence measurements revealed the tetragonal (TP)-cubic (CP) phase transition to occur between 170-200 K and the orthorhombic (OP)-TP transition to occur between 80 and 100 K. (B) Comparison of the dependence of the Gibbs free energy on temperature for different polymorphs of a compound in the bulk and nanocrystal state. Reproduced with permission from ref. 130. Copyright 2019, American Chemical Society.

temperature for the different bulk and nanoconfined phases that is consistent with our experimental observations.

Examining nanoconfined CsPbI3 crystallization as a function of AAO pore diameter, Ma and coworkers found that the black phase of CsPbI₃ (reported as the α phase but weak superstructure peaks in the XRD patterns likely indicate the γ phase was present) became thermodynamically stable within 41 nm diameter pores or smaller. 151 DFT calculations suggested that stabilization of the black CsPbI3 phase at room temperature was due both to increased surface area of nanoconfined crystals and strain introduced at the pore walls. By tuning the processing conditions, we discovered that γ-CsPbI₃ could be stabilized in AAO pores as large as 250 nm. ¹⁵² Using wet annealing method in which spun cast samples with solvent remained were immediately annealed at a minimum temperature 373 K, γ-CsPbI₃ was formed within the AAO nanopores. Once formed, nanoconfined γ-CsPbI₃ was found to be stable in a range of 4-610 K and for a period of at least 120 days of storage in air. Using chemical vapor deposition to infiltrate the 250 nm pores of AAO scaffolds, Waleed and coworkers recently reported stabilization of α-CsPbI₃ at room temperature. ¹⁵³ In comparison, CsPbI₃ deposited via spin coating in the same nanopores was found to adopt the orthorhombic γ phase, indicating that processing conditions play a crucial role in determining final crystal structures. Nanoconfined stabilization of the metastable α phase of FAPbI₃ from conversion to the thermodynamically stable δ -phase at room temperature has also been reported, 154 although shifting of polymorph transitions temperatures was not reported.

Influence of pressure on phase transitions

Solid state phase transitions are a function of both temperature and pressure. Enabled by advances in synchrotron-based pressurized configurations for high-resolution diffraction measurements, 155 pressured-induced phase transitions in MHPs have been intensely studied in recent years. Compared to other perovskites, MHPs exhibit high compressibility and pressure-dependent tunability of optoelectronic properties. 156,157 Several reviews have already been published on pressure-induced phase transitions in MHPs and the role of pressure in modulating electronic landscapes, 155,157-159 so we will only provide a brief overview here.

In general, MHPs undergo similar phase transitions upon pressurization in the range of atmospheric to 50 GPa. 159 At lower pressures, structural changes are characterized by compression of Pb-I bonds and/or octahedral tilting, as displayed in Fig. 10.160 At higher pressures, MHPs amorphize, a process which is reversible upon pressure reduction. MAPbCl₃, for example, undergoes a cubic-to-cubic isostructural transition upon increasing the pressure from atmospheric pressure to 0.8 GPa due to a continuous decrease in the Pb-Cl bond length. 161 Increasing the pressure further to 2.4 GPa was found to induce octahedral tilting such that the crystal structure transitions to an orthorhombic unit cell with Pnma symmetry. Fig. 11 summarizes the pressure-dependent phase behavior of 12 perovskites, including CsPbCl $_3$, 162 CsPbBr $_3$, 163 CsPbI $_3$, 164 CsSnBr $_3$, 165 MAPbCl $_3$, 161 MAPbBr $_3$, 166 MAPbI $_3$, 167 MASnCl $_3$, 161 MASnBr $_3$, 165

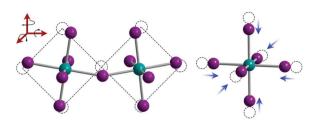


Fig. 10 Pressure-induced rotation of PbI₆ octahedra (left) and Pb-I bond compression (right). Reproduced with permission from ref. 160. Copyright 2017, American Chemical Society

MASnI₃, ¹⁶⁸ FAPbBr₃, ¹⁶⁹ and FAPbI₃. ¹⁷⁰ CsPbCl₃, CsPbBr₃ and MaPbCl₃ all exhibit an isostructural phase transition in which the crystal symmetry is retained upon increasing pressure via symmetric compression of the M-X bond. We note that there are some discrepancies in reported pressure-induced phase transitions, for example in widely studied MAPbI₃, ^{167,171-174} which partially result from differing ambient pressure phase assignments.

In the opposite direction, lowering the pressure can also impact solid state phase transitions in MHPs. Decreasing the pressure to 10^{-3} mbar was recently found to shift the tetragonal to cubic phase transition temperature of MAPbI₃ higher by 40 °C, from \sim 323-363 K.¹⁷⁵ Furthermore, the tetragonal and cubic phases were observed to coexist until a temperature of 140 °C was reached. This shift was attributed to a reduction in temperature-induced volume expansion rate of MAPbI3 upon pressure reduction.

Promisingly, recent reports have found that some pressureinduced changes can be retained upon cycling back to ambient pressure. For FAPbI₃ exposed to pressures of 7.1 GPa, it was found that unit cell parameters at ambient pressure after compression were smaller compared to before compression, as evidenced by a shift in diffraction peaks to higher 2-theta values. 176 This structural change resulted in a narrower bandgap of previously compressed versus uncompressed FAPbI3. Recently, quenching MAPbI₃ from a high pressure state was found to result in retention of bandgap reduction upon pressurization.177 Partial retention of pressure-induced structural changes has also been reported for 2D perovskites. 178,179

Polymorph-dependent optoelectronic properties

Among compounds adopting the AMX₃ perovskite crystal structure, those incorporating halide ions in the X position have sufficiently small bandgaps, from $\sim 1.2-1.7$ eV, 42 to absorb visible light and transport charge carriers for use in solar cells and other optoelectronic devices. First principle calculations have found that the bandgap is determined by the metal-halide framework, with A-site cations contributing minimally to the band structure. 180,181 The valence band maximum comprises antibonding metal s and halogen p orbitals, while the conduction band minimum primarily comprises empty metal p orbitals. 181,182 The band structure of MHPs is sensitive to slight changes in the M-X framework due to lattice contractions and octahedral tilting. 183 Thus while not directly contributing to the band structure, A-site cations indirectly influence band structures through altering the M-X framework.

These trends are exemplified in a recent comparison of the temperature-dependent optical bandgaps of CsPbBr₃, MAPbBr₃, and FAPbBr₃ single crystals. 184 As observed from Fig. 12, optical bandgaps decrease with increasing tolerance factors of 0.92, 1.04, and 1.14 for CsPbBr₃, MAPbBr₃, and FaPbBr₃, respectively. 15,185,186 Overall, the bandgaps of halide perovskites are strongly correlated to the M-X-M bond angle, 187,188 which can range from 150–180°, as described in detail in a 2015 account by Stoumpos and Kanatzidis.19 Upon bending, interactions between M p orbitals and X p antibonding orbitals increases, resulting in an increase of the bandgap. 183 Significant efforts in the field have focused on exploiting the relationship between A-site cation size, the extent of M-X-M bond angle distortion, and the overall bandgap of the system through the engineering of mixed composition MHPs. ^{189–195} For a series of MAPb($I_{1-x}Br_x$)₃ compounds, for example, Noh and coworkers reported the ability to continuously tune the bandgap from 1.6 to 2.3 eV by increasing the value of x from 0 to 1. Fig. 13 displays the bandgaps of six different MHPs as a function of the X-M-X

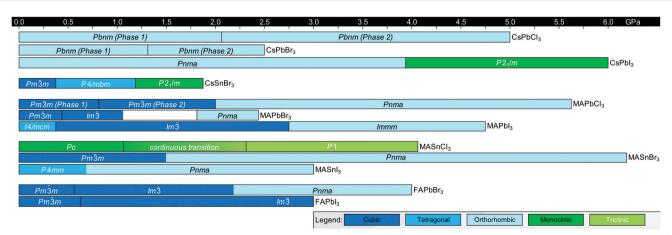


Fig. 11 Summary of room-temperature phase transition pressures in halide perovskites as a function of composition reported in the literature. 162–170 The white section in the data for MAPbBr₃ indicates a region where data was not available.

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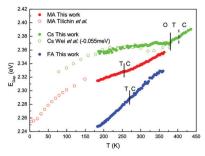


Fig. 12 Comparison of optical bandgap of MAPbBr₃, CsPbBr₃, and FAPbBr₃ single crystals. Cubic, tetragonal and orthorhombic phases are indicated by C, T, and O labels, respectively. Reproduced with permission from ref. 184. Copyright 2020, American Chemical Society

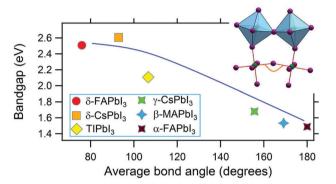


Fig. 13 Graph of the bandgap versus M-X-M bond angle in MHPs. Reproduced with permission from ref. 157. Copyright 2019, Elsevier.

bond angle. 157 As observed from the figure, the bandgap decreases as this angle approaches 180°.

Following this trend of widening bandgaps with increased octahedral tilting among different MHP compounds, one might expect that within a single compound, the bandgap should decrease with increasing temperatures as MHPs transition from low-symmetry to high-symmetry structures. It has been demonstrated experimentally, however, that the bandgaps of MHPs increase with increasing temperature. 197 Indeed, the temperature dependence of the optical bandgaps of CsPbBr₃, MAPbBr3, and FAPbBr3 in Fig. 12 all display a positive temperature coefficient, with monotonically increasing bandgaps with increasing temperature. For widely studied MAPbI₃, a blue shift in the bandgap with increasing across the tetragonal-cubic transition has been recorded via temperature-dependent photoluminescence, 198-203 spectroscopic ellipsometry, 204 and vis-NIR absorption spectroscopy205 even though the nominal Pb-I-Pb angle increases from 165.3 to 180° across the tetragonal-cubic transition.

Through a combination of temperature-dependent UV-vis absorption and ab initio simulations, Quarti and coworkers found that MAPbI3 monotonically blue-shifts through this transition due to large fluctuations in the M-X framework on the picosecond time scale that increase with increasing temperature.²⁰⁶ At temperatures above tetragonal-to-cubic phase transition, MAPbI₃ was found to adopt the cubic structure on average, but in fact fluctuates between highly distorted configurations, resulting in an increase in the bandgap compared to the tetragonal phase. Computational modeling has attributed the increase in bandgap upon temperature increase to electron-phonon coupling and thermal expansion. 197,207,208

Interestingly, we recently observed a ~ 30 meV red shift in the PL spectrum of MAPbI3 across the tetragonal-cubic transition when MAPbI₃ crystals were nanoconfined in the $\sim 20-250$ diameter pores of AAO (Fig. 9), 150 opposite of the trend observed in bulk crystals. By increasing the surface area-to-volume ratio of these crystals, nanoconfinement shifted the relative Gibbs free energies of the α , β , and γ phases such that the α - β transition temperature was lowered from ~ 330 K to between 170-200 K. At these lower temperatures, thermal fluctuations in the M-X framework are likely suppressed, resulting in a narrowing of the bandgap upon transitioning to the cubic phase with I-Pb-I angles of 180°. This narrowing of the bandgap of α-MAPbI₃ compared to β-MAPbI₃ is predicted when considering static structures. 27,209,210

Examining pressure-induced changes in optoelectronic properties of MHPs has also advanced a fundamental understanding of the relationship between molecular structure and band structure. Pressure-induced compression of M-X-M bonds results in a decrease in the bandgap due to increased overlap between Pb 6s and I 5p orbitals, increasing the valence band energy and decreasing the overall bandgap. 171,211 A linear positive correlation between Pb-I-Pb bond length and the bandgap, for example, has been reported for MAPbI₃. 172 On the other hand, octahedral tilting reduces this overlap, resulting in an increase in bandgap at higher pressures. 159,167,172

Beyond the bandgap, the impact of polymorphic phase transitions on other optoelectronic properties varies for different compositions and temperature ranges. For CsPbI3, CsSnI3, and FAPbI₃, polymorph transitions to their non-perovskite δ -phases destroy device performance as these phases are insulating. For transitions between α , β , and γ phases, the impact on optoelectronic properties is less catastrophic, but can be detrimental nonetheless. While many questions still remain about optoelectronic processes in these materials, 212,213 recent temperaturedependent studies have begun to track photophysical processes across polymorph transitions, particularly in widely studied MAPbI₃.

In 2015, for example, Milot and coworkers used temperaturedependent time-resolved PL and THz conductivity measurements to examine charge recombination in MAPbI3. 11 The rate of monomolecular recombination was observed to increase smoothly in the range of 8-340 K. On the other hand, the temperature dependence of bimolecular and Auger recombination rates were found to depend on the polymorph, as displayed in Fig. 14. In particular, Auger recombination rates significantly increased upon transition from the tetragonal to orthorhombic phase. In a later study, pulse-radiolysis time-resolved microwave conductivity experiments on MAPbI₃ and MAPbBr₃ further revealed abrupt increases in charge carrier mobilities and lifetimes when crystals transitioned from the tetragonal to the orthorhombic phase.²¹⁴ While MA⁺ cations can rotate in the tetragonal phase, their orientation becomes fixed in the orthorhombic phase to form

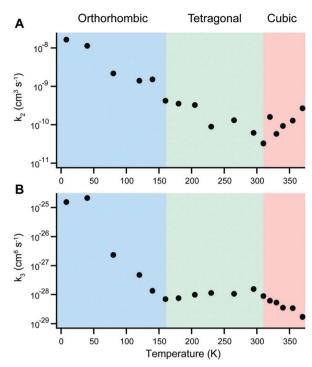


Fig. 14 Temperature dependence of (A) bimolecular and (B) Auger recombination rates for MAPbl₃. Reproduced with permission from ref. 11. Copyright 2015, Wiley.

charged domains that reduce charge recombination. This ordering of the MA⁺ cation was found to result in smaller dielectric constants and larger exciton binding energies compared to the tetragonal phase, contributing to a faster rate of ground state bleaching recovery in the former phase. 215 In contrast, no abrupt changes across the tetragonal-orthorhombic phase transition for FAPbI₃ was observed due to the fact that FA⁺ retain rotational freedom in the low-temperature orthorhombic phase.²¹⁶

Because the MA⁺ cations can rotate in all directions in both the tetragonal and cubic phases of MAPbI3, their role in modulating photophysical processes across the tetragonalcubic transition may be less pronounced compared to the orthorhombic-tetragonal transition. Examining temperaturedependent dielectric properties of MAPbI3 using microwave conductance experiments found that cation rotation was too slow to influence charge carrier mobilities and lifetimes in the range of 160-300 K.217 Recent temperature-dependent photoconductivity experiments on FAPbBr3, on the other hand, revealed multiple discontinuities in the peak intensities and peak locations, only two of which corresponded to crystallographic phase changes from the cubic to tetragonal phase at 266 K and the tetragonal to orthorhombic phase at 153 K.²¹⁸ Fig. 15 displays the photoconductivity characteristics, heat capacity, and mean square displacement (MSD) of hydrogen atoms as a function of temperature. Five distinct discontinuities can be observed in the temperature-dependent excitonic peak center, represented by gray circles in Fig. 15a. These five transitions were all found to correspond to discontinuities in the temperature-dependent MSD of hydrogen, as measured by

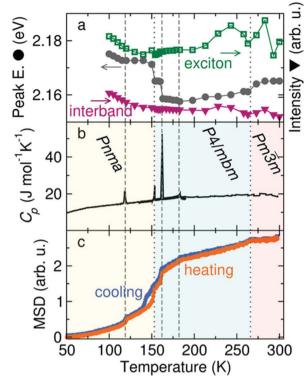


Fig. 15 (a) Plot of the excitonic peak intensity (green squares), excitonic peak center (gray circles) and interband intensity (purple triangles) extracted from photoconductivity data as a function of excitation wavelength and temperature for FAPbBr₃. (b) Differential scanning calorimetry scan for FAPbBr₃, with crystallographic phases at different temperatures labelled. (c) Mean squared displacement of hydrogen upon heating and cooling of FAPbBr₃ measured via elastic neutron scattering. Dotted and dashed lines represent discontinuities in the spectra that do and do not correspond to solid-state phase transitions, respectively. Reproduced with permission from ref. 218. Copyright 2019, American Chemical Society.

elastic neutron scattering (Fig. 15c). These results suggest that FA+ rotational dynamics may induce local temperaturedependent distortions through dynamic coupling to the M-X framework.

One of the main applications for MHPs is in emerging photovoltaics to harvest renewable solar energy. Depending on the location and specific use, solar panels can be exposed to a wide range of temperatures, inducing polymorph transitions in the MHP active layer during device operation. Monitoring structural phase transitions in MAPbI₃ active layers during solar cell operation, Zhang and coworkers observed a drop in photocurrent from 20 mA cm⁻² at 240 K to 1 mA cm⁻² at 140 K.²¹⁹ This large difference in photocurrent was attributed to a phase transition from the tetragonal to orthorhombic phase upon lowering the temperature. In these devices, power conversion efficiencies of <0.1% were measured at temperatures below 160 K when the orthorhombic phase was present, compared to a maximum efficiency of 16% at 300 K for tetragonal MAPbI₃.²¹⁹ Recent time-of-flight and electrical conductivity measurements on MAPbI3 solar cells have further revealed a balancing of electron and hole mobilities at the tetragonal-cubic phase transition.²²⁰

Conclusion

Review

The chemical diversity and flexibility of MHPs has given rise to myriad possibilities of structures and compositions for use in optoelectronic devices. Systematic experimental and computational studies of the relationships between A-site, M-site and X-site ion sizes and the crystal structure of MHPs, as well as temperature-dependent measurements of these crystals via high-resolution X-ray and neutron diffraction, have provided a comprehensive understanding of the factors governing polymorphism in this class of materials. In general, MHPs with tolerance factors closer to 1 will adopt a cubic structure at room temperature, whereas octahedral tilting in MHPs with tolerance factors farther from 1 result in orthorhombic or tetragonal crystal structures under ambient conditions. MHPs transition from lower to higher symmetry with increasing temperature, but higher to lower symmetry with increasing pressure. More complex and less understood are how slight changes in bond lengths and angles, as well as A-site cation dynamics, affect the electronic structure and optoelectronic processes, such as exciton recombination and charge transport. Looking forward, optoelectronic characterization across phase transitions will provide further mechanistic insights into the interplay between charges and the metal halide framework. The role of strategies to shift solid-state phase transition temperatures, for example via size reduction or nanoconfinement, and to kinetically trap metastable phases, such as through pressure cycling, have the potential to improve both the performance and reliability of MHP-based devices. Looking towards future commercialization, suppression of polymorph transitions that negatively impact device performance will be necessary, and research on structure-property-function relationships across these solid-state transitions will undoubtedly play a critical role in the advancement of MHPs for optoelectronic applications.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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