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The design of second generation MOP-phosponites: efficient chiral hydrosilylation of functionalised styrenes†

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A series of enantiopure MOP-phosponite ligands, with tailored steric profiles, have been synthesised and are proven to be very successful in high-yielding, regio- and enantioselective catalytic hydrosilylation reactions of substituted styrenes, affording important chiral secondary alcohols.

Chiral mono- and bidentate phosphorus ligands have achieved considerable acclaim in a variety of transition metal-catalysed asymmetric transformations.¹ Atropisomeric monophosphines with a binaphthyl backbone² have proven to be particularly successful in chiral hydrosilylations of alkenes.³

We recently reported a new class of MOP-phosponite ligands (**1a** and **1b**), based on a MOP/XuPhos-type hybrid (Chart 1).⁴ Chiral phosponites are ligands with good π -acceptor properties and are effective in a range of asymmetric hydrocyanations,⁵ hydroformylations⁶ and hydrogenations.⁷ Ligands **1a** and **1b** possess a phosphorus donor bound directly to a binaphthyl backbone – the **a** series has a hydrogen in the 2' position of the lower naphthyl ring whilst a methoxy group occupies the site in the **b** series (Chart 1). This difference has profound implications for the catalytic performance of the

parent (*S*)-H-MOP and (*R*)-MeO-MOP ligands, and we wished to examine this effect for phosponite derivatives. Importantly, we also wanted to examine the effect of incorporating a second chiral moiety, which we achieved by appending either (*R*)- or (*S*)-binol to isolate the diastereomeric pairs (*S*,*R*_b)-**1a**/*(S,S*_b)-**1a** and (*R*,*R*_b)-**1b**/*(R,S*_b)-**1b**. By investigating the coordination chemistry and catalytic performance of the four ligand pairs, we were able to demonstrate that subtle differences in the position of the palladium atom relative to the lower naphthyl ring in these complexes appears to have important consequences for chiral induction in the catalytic hydrosilylation of styrene.⁴ These results led us to question whether the large chiral binol group was in fact necessary for effective asymmetric catalysis.

In this study we sought to investigate the effect of replacing the binol moiety of **1a** and **1b** with achiral ancillary aryloxides, in order to establish if the enantioselectivities remain high by virtue of the phosphorus-bonded binaphthyl backbone, and thus eliminate the requirement for a second chiral centre. To the best of our knowledge this has not been demonstrated for MOP phosponites, which are likely to effectively catalyse alternate substrates than their phosphine cousins, by virtue of their differing electronic properties. Herein we describe the synthesis of the phenoxy derivatives (*S*)-**3a** and (*R*)-**3b**, which allowed us to probe the effect of removing the chelating element of the added aryoxide, before describing how the sensitivity of these ligands to hydrolysis demanded the design of next generation biphenoxy-based MOP-phosponites (*S*)-**4a** and (*R*)-**4b**, which were further improved upon to yield the user-friendly, chiral phosponites (*S*)-**5a** and (*R*)-**5b**, that perform better than the corresponding MOP phosphines in the hydrosilylation of functionalised styrenes (Chart 2 and Table 2).

The six new MOP-phosponite ligands were synthesised following a procedure we have developed, starting with the air-stable, chiral primary phosphine precursors (*S*)-**1a** and (*R*)-**1b**.⁸ In a two-step, one-pot reaction, the primary phosphine was treated with phosphorus pentachloride, quantitatively yielding the dichlorophosphines (*S*)-**2a** and (*R*)-**2b** (Scheme 1).⁴ Our initial targets (*S*)-**3a** and (*R*)-**3b** were synthesised by employing

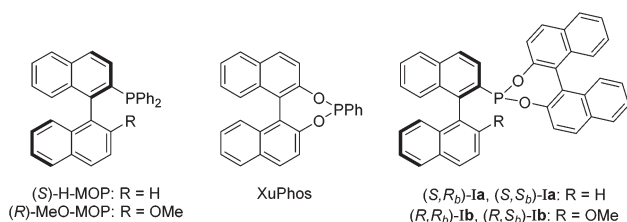


Chart 1 A selection of binaphthyl-containing monophosphorus-donor ligands.

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† Electronic supplementary information (ESI) available: Full experimental and crystallographic data. CCDC 1416469 (*R*)-**3b**, 1416470 (*R*)-**4b**, 1416471 (*S*)-**6a** and 1416472 (*S*)-**7a**. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5dt04475h



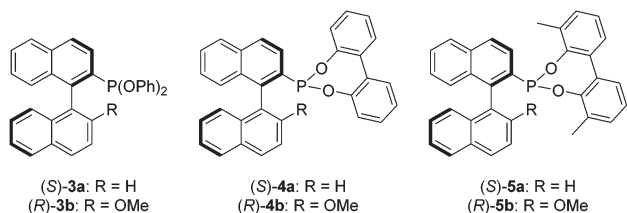
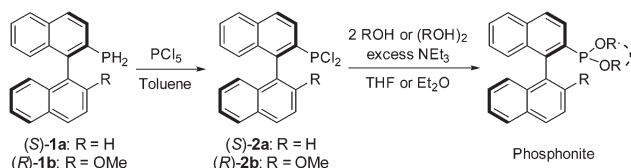


Chart 2 Phosphonite ligands (S)-3a, (R)-3b, (S)-4a, (R)-4b, (S)-5a and (R)-5b.



Scheme 1 The synthesis of the phosphonite ligands from (S)-1a and (R)-1b.

two equivalents of phenol, however both syntheses proved to be temperamental as a result of their sensitivity to hydrolysis, with (S)-3a particularly low-yielding. Therefore we sought to replace the two phenol groups with the achiral, chelating 2,2'-biphenol, in order to ascertain whether we could inhibit this decomposition. This approach led to the successful preparation of the enantiopure compounds (S)-4a and (R)-4b; the ligands can be stored under nitrogen for a period of weeks with only minimal decomposition. However, during our initial investigations into the coordination chemistry of these ligands, we noticed slow decomposition in solution and their performance in catalytic hydrosilylation required improvement (*vide infra*). Thus, in order to push the design concept still further and optimise both ligand stability and catalytic capability, we synthesised the bulky *ortho*-methyl substituted starting material 2,2'-dihydroxy-3,3'-dimethyl-1,1'-biphenyl (ESI†), and used it to prepare the ligands (S)-5a and (R)-5b which possess a more imposing steric profile. Both phosphonites were isolated in high yield, and are purified by passing through a plug of alumina in air. They can be stored under nitrogen without decomposition over several months, and their coordination complexes also exhibit better stability.

The phosphonites were fully characterised by multinuclear NMR spectroscopy and High Resolution Mass Spectrometry. Their $^{31}\text{P}\{^1\text{H}\}$ spectra show characteristic singlet peaks; δ (ppm) = (S)-3a (154.8), (R)-3b (155.6), (S)-4a (177.7), (R)-4b (180.0), (S)-5a (172.4) and (R)-5b (174.2). The resonances for the phenoxy-derived phosphonites have a notable upfield shift compared to the biphenoxyl analogues. Single crystals of (R)-3b and (R)-4b suitable for X-ray crystallographic analysis were obtained (Fig. 1 and 2); there are two crystallographically independent molecules in the asymmetric unit of (R)-3b. Density Functional Theory calculations at the B3LYP/6-31G* level of theory (ESI†) revealed that the phosphonites have

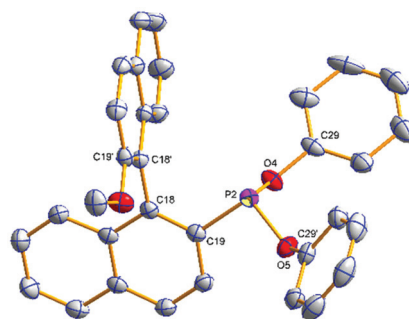


Fig. 1 Molecular structure of (R)-3b. Hydrogen atoms have been omitted for clarity. Selected average bond distances (Å) and angles (°): P2–C19 1.830(2), P2–O4 1.6461(17), P2–O5 1.6619(18), O4–C29 1.402(3), O5–C29' 1.391(3); O4–P2–O5 99.13(9), O4–P2–C19 93.63(9), O5–P2–C19 97.85(10), C19–C18–C18'–C19' –86.5(3).

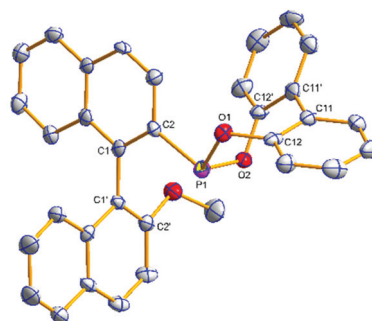


Fig. 2 Molecular structure of (R)-4b. Hydrogen atoms have been omitted for clarity. Selected average bond distances (Å) and angles (°): P1–C2 1.812(7), P1–O1 1.658(4), P1–O2 1.657(4), O1–C12 1.411(7), O2–C12' 1.396(7), C11–C11' 1.482(9); O1–P1–O2 99.5(2), O1–P1–C2 96.6(2), O2–P1–C2 100.3(3), C2–C1–C1'–C2' –86.5(7), C12–C11–C11'–C12' –48.6(8).

lower HOMO (Highest Occupied Molecular Orbital) and LUMO (Lowest Unoccupied Molecular Orbital) energies than their aryl counterparts, and have a lower phosphorus contribution to the HOMO, implying the phosphonites are weaker P-ligand σ -donors. To study the coordination chemistry of our ligands with Pd(II), we reacted them with $[\text{Pd}(\eta^3\text{-C}_4\text{H}_7\text{-Cl})_2]$ in a 1 : 1, P : Pd ratio, to quantitatively synthesise the complexes $[(\eta^3\text{-C}_4\text{H}_7)\text{PdCl}(\text{L}^{\text{P}})]$.

Single crystals of $[(\eta^3\text{-C}_4\text{H}_7)\text{PdCl}(\text{S})\text{-}(\mathbf{4a})]$ ((S)-6a) and $[(\eta^3\text{-C}_4\text{H}_7)\text{PdCl}(\text{S})\text{-}(\mathbf{5a})]$ ((S)-7a) suitable for X-ray analysis were grown (Fig. 3 and 4); the asymmetric unit of (S)-7a comprises two molecules in different conformations (the second structure and an overlay is provided in the ESI†) and there is twofold disorder of the methallyl ligand and the palladium atom in one of these independent. The phosphonites show monodentate coordination through the phosphorus, and there is a *pseudo*-square-planar geometry around the palladium.

The allyl carbons *trans* to phosphorus exhibit a longer Pd–C bond length (2.195, 2.206 Å) than the allyl carbons *cis* to the phosphorus (2.096, 2.095 Å), which is consistent with the



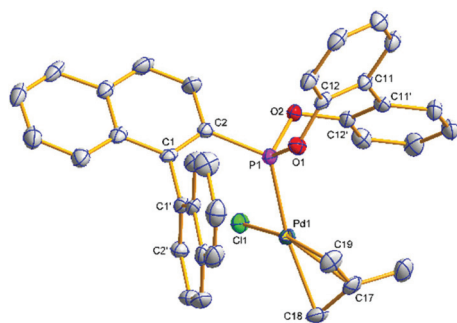


Fig. 3 Molecular structure of (*S*)-**6a**. Hydrogen atoms have been omitted for clarity. Selected average bond distances (Å) and angles (°): Pd1–P1 2.2478(10), Pd1–C11 2.3583(9), Pd1–C17 2.162(4), Pd1–C18 2.195(4), Pd1–C19 2.096(4), P1–C2 1.815(4), P1–O1 1.619(3), P1–O2 1.610(3), C11–C11' 1.477(5); P1–Pd1–C11 91.55(4), P1–Pd1–C18 165.51(11), P1–Pd1–C19 98.27(13), C18–Pd1–C19 67.46(16), C18–C17–C19 114.6(4), C2–C1–C1'–C2' –80.8(5), C12–C11–C11'–C12' 42.3(6).

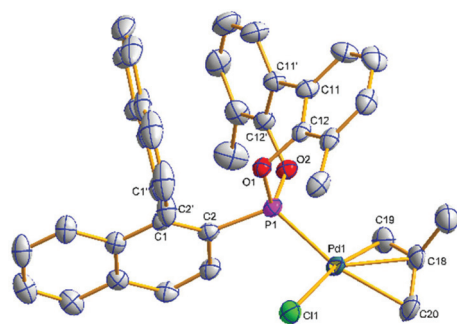


Fig. 4 Molecular structure of (*S*)-**7a**. Hydrogen atoms have been omitted for clarity. Selected average bond distances (Å) and angles (°): Pd1–P1 2.2368(11), Pd1–C11 2.3739(11), Pd1–C18 2.152(4), Pd1–C19 2.095(4), Pd1–C20 2.206(4), P1–C2 1.804(4), P1–O1 1.610(3), P1–O2 1.622(3), C11–C11' 1.485(6); P1–Pd1–C11 94.59(4), P1–Pd1–C19 96.50(15), P1–Pd1–C20 162.95(15), C19–Pd1–C20 66.8(2), C19–C18–C20 117.4(5), C2–C1–C1'–C2' –81.9(6), C12–C11–C11'–C12' –46.0(7).

stronger *trans* influence of the phosphonite compared to the chloride ligand. In (*S*)-**6a** the lower naphthyl fragment of the binaphthyl backbone is face-to-face with the palladium centre (Fig. 3). Neither of the two independent of (*S*)-**7a** display this feature, with the lower naphthyl group being orientated away from the palladium centre and facing the biphenyl moiety (Fig. 4 and ESI†). In the two independent of (*S*)-**7a** the torsion angle of the dimethyl substituted biphenyl moiety is of opposite sign, and also when comparing the biphenyl moiety in (*R*)-**4b** and (*S*)-**6a** (Fig. 2 and 3), implying no restriction to rotation about the C11–C11' bond.

The $^{31}\text{P}\{^1\text{H}\}$ NMR peaks of (*S*)-**6a** and (*S*)-**7a** experience slight shifts compared to the free ligands, and show the presence of two independent resonances due to the two isomers formed; δ (ppm) = (*S*)-**6a** (172.0 and 173.6) and (*S*)-**7a** (172.9 and 177.6). The isomers are a result of the rotation of the allyl moiety, *via* a selective *syn/anti* exchange of the allyl

protons *cis* to the phosphorus, by a $\eta^3\text{-}\eta^1\text{-}\eta^3$ mechanism, due to the stronger *trans* effect of the phosphonite donor ligand.^{4,9} The rapid exchange process resulted in broadened peaks in the NMR spectra at room temperature; cooling of CD_2Cl_2 solutions of the complexes to -20°C sharpened the resonances and allowed for full characterisation of the methallyl ligands in both isomers (ESI†). Reaction of (*S*)-**7a** with NaBAR^{F} resulted in loss of NaCl and the formation of $[(\eta^3\text{-C}_4\text{H}_7)\text{Pd}(\text{S})\text{-}(\mathbf{5a})]\text{BAR}^{\text{F}}$ (**8a**), where (*S*)-**5a** acts as a chelating P,C- π -donor. The upfield $^{13}\text{C}\{^1\text{H}\}$ NMR coordination chemical shifts for both C1' and C2' suggest a η^2 -binding mode (Fig. 5), which is in agreement with the results of an NMR study reported by Pregosin and co-workers.⁹

As discussed, MOP phosphine ligands are known to give high enantioselectivity in the palladium-catalysed asymmetric hydrosilylation of alkenes, particularly styrene, to give chiral secondary alcohols (Scheme 2).³ To gain an insight into how the different stereoelectronic profiles of our phosphonites impacts upon their catalytic activity in the same transformation, we prepared catalysts by reacting each phosphonite with $[\text{Pd}(\eta^3\text{-C}_3\text{H}_5)\text{Cl}]_2$ (Table 1). We chose to test our phosphonites against the well-known H-MOP and MeO-MOP phosphines (the latter is a commercial compound), employing P : Pd ratios of 1 : 1 and 2 : 1 at room temperature – full conversion was obtained in all cases. We noted a general increase in enantioselectivity and reaction rate in the order (*S*)-**5a**/*(R)*-**5b** > (*S*)-**4a**/*(R)*-**4b** > (*S*)-**3a**/*(R)*-**3b**, with the introduction of the biphenyl moiety, and subsequently the methyl groups, markedly improving the ligand performance. Phosphonite (*S*)-**5a** gave excellent enantioselectivity for (*R*)-1-phenylethanol (Table 1, 95%, entry 18), which is slightly higher than that for the (*S*)-H-MOP phosphine (Table 1, 94%, entry 4), although the latter reaction reached conversion more quickly. We also tested the H-MOP phosphonite ligands (*S,R_b*)-**1a** and (*S,S_b*)-**1a** at 0°C , these ligands are far less selective than (*S*)-**5a**.⁴

Interestingly, although the methoxy-substituted ligands (*R*)-MeO-MOP phosphine and (*R*)-**5b** performed poorly compared

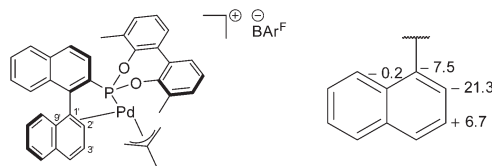
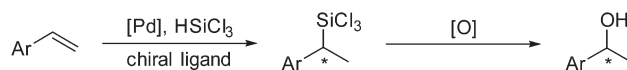


Fig. 5 Proposed structure of (*S*)-**8a** (left) and a fragment showing the $^{13}\text{C}\{^1\text{H}\}$ NMR coordination chemical shift (ppm) between the major isomers of (*S*)-**7a** and (*S*)-**8a** (right).



Scheme 2 The synthesis of chiral secondary alcohols via the palladium catalysed asymmetric hydrosilylation and subsequent oxidation of styrene derivatives.



Table 1 Catalytic results from the asymmetric hydrosilylation of styrene^a

	L	P : Pd	T	Time ^b	Conv. ^c	ee ^d
1	(S)-H-MOP	1 : 1	rt	1 min	>99%	74 (R)
2	(S)-H-MOP	1 : 1	0 °C	4 h	>99%	93 (R)
3	(S)-H-MOP	2 : 1	rt	1 h	>99%	75 (R)
4	(S)-H-MOP	2 : 1	0 °C	12 h	>99%	94 (R)
5	(R)-MeO-MOP	1 : 1	rt	10 min	>99%	22 (R)
6	(R)-MeO-MOP	2 : 1	rt	1 h	>99%	20 (R)
7	(S)-3a	1 : 1	rt	24 h	>99%	63 (R)
8	(S)-3a	2 : 1	rt	24 h	>99%	73 (R)
9	(R)-3b	1 : 1	rt	24 h	>99%	7 (S)
10	(R)-3b	2 : 1	rt	24 h	>99%	1 (S)
11	(S)-4a	1 : 1	rt	6 min	>99%	73 (R)
12	(S)-4a	2 : 1	rt	1 h	>99%	79 (R)
13	(R)-4b	1 : 1	rt	10 min	>99%	23 (R)
14	(R)-4b	2 : 1	rt	1 h	>99%	27 (R)
15	(S)-5a	1 : 1	rt	2 min	>99%	79 (R)
16	(S)-5a	1 : 1	0 °C	5 h	>99%	92 (R)
17	(S)-5a	2 : 1	rt	2 h	>99%	80 (R)
18	(S)-5a	2 : 1	0 °C	72 h	>99%	95 (R)
19	(R)-5b	1 : 1	rt	1 h	>99%	45 (R)
20	(R)-5b	2 : 1	rt	70 min	>99%	51 (R)
21	(S,R _b)-1a	1 : 1	0 °C	168 h	88% ^e	4 (S)
22	(S,R _b)-1a	2 : 1	0 °C	168 h	>99%	60 (R)
23	(S,S _b)-1a	1 : 1	0 °C	72 h	>99%	59 (R)
24	(S,S _b)-1a	2 : 1	0 °C	168 h	22% ^e	77 (R)

^a The catalyst was generated *in situ* from the ligand (0.25 or 0.50 mol%) and [Pd(η^3 -C₃H₅)Cl]₂ (0.125 mol%), and reacted with styrene (10 mmol) and trichlorosilane (12 mmol). ^b Time taken for reaction to reach completion. ^c Determined by ¹H NMR spectroscopy. ^d % ee determined by chiral GC; absolute configuration assigned by comparison of the sign of optical rotation to literature (ESI). ^e Reaction stopped.

to their H-substituted counterparts – an established trend for this ligand family¹⁰ – the phosphonite ligand gave much higher enantioselectivities than the corresponding phosphine (51% *versus* 20%, Table 1, entries 20 and 6), demonstrating further the potential these derivatives have to improve existing toolkits.

It is notable that at room temperature, the hydrosilylation occurs as a highly exothermic, quite violent reaction which is accompanied by a spontaneous colour change to deep black; similar observations with arylmonophosphinoferrrocenes have also been made.¹¹ When a 2 : 1, P : Pd ratio was used there was often a significant delay in the time taken to observe this phenomenon (from 1–8 h, Table 1). However these exothermic colour changes were not observed when the external temperature of the flask was kept at 0 °C. Consequently, the catalysis performed at 0 °C using both (S)-H-MOP and the sterically bulky (S)-5a resulted in longer reaction times, however this was offset by an increased ee of the product – phosphonite (S)-5a gave the best ee with a value of 95%.

The final objective in this study was to extend our styrene screening regimen to substrates that MOP phosphines struggle to catalyse, or confer only low or moderate enantioselectivity in the product, with the aim that the unique stereoelectronic properties of the phosphonites would allow for improvements to be made. As such we extended our screening with (S)-5a to include the *para*-substituted styrenes listed in Table 2. Intro-

Table 2 Asymmetric hydrosilylation of heteroatom-substituted styrenes^a

	L	Ar	T	Time ^b	Conv. ^c	ee ^d
1	(S)-H-MOP	4-ClC ₆ H ₄	rt	1 h	>99%	78 (R)
2	(S)-H-MOP	4-ClC ₆ H ₄	0 °C	24 h	>99%	94 (R)
3	(S)-5a	4-ClC ₆ H ₄	rt	18 h	>99%	86 (R)
4	(S)-5a	4-ClC ₆ H ₄	0 °C	336 h	<5% ^e	—
5	(S)-H-MOP	4-MeOC ₆ H ₄	rt	10 min	>99%	35 (R)
6	(S)-H-MOP	4-MeOC ₆ H ₄	0 °C	12 h	>99%	58 (R)
7	(S)-5a	4-MeOC ₆ H ₄	rt	70 min	>99%	78 (R)
8	(S)-5a	4-MeOC ₆ H ₄	0 °C	240 h	>99%	85 (R)

^a The catalyst was generated *in situ* from the ligand (0.50 mol%) and [Pd(η^3 -C₃H₅)Cl]₂ (0.125 mol%) and reacted with styrene (10 mmol) and trichlorosilane (12 mmol). ^b Time taken for reaction to reach completion. ^c Determined by ¹H NMR spectroscopy. ^d % ee determined by chiral GC; absolute configuration assigned by comparison of the sign of optical rotation to literature (ESI). ^e Reaction stopped.

ducing an electron withdrawing chloride group resulted in much longer reaction times, however full conversion and a high ee, better than (S)-H-MOP, were again obtained (Table 2, entries 1 and 3; note entry 2 for improving the ee generated by the latter). The chloro-substituted chiral secondary alcohol product was used to synthesise the antihistamine Clemastine enantioselectively.¹² More significantly, when an electron donating 4-methoxy substituent was incorporated onto the styrene precursor, (S)-5a was found to be a remarkably superior ligand, generating a much higher ee than (S)-H-MOP, albeit more slowly (85% *versus* 58%, Table 2, entries 8 and 6). It is of note that the product alcohol from this reaction was used to prepare antagonists of lysophosphatidic receptors, which may have implications in the treatment of lung fibrosis, liver disease and cancer.¹³ Our future work will seek to extend the substrate scope to styrenes bearing other important functional groups.

In summary, we have synthesised a series of novel MOP-phosphonite ligands from air-stable, chiral primary phosphine precursors, and have demonstrated the ease with which both the sterics and the hydrolytic nature of the ligands can be tuned, by modifying the aryl groups of the aryloxy moiety. In comparison to our previous report on **1a/1b**, a major finding here is that a second chiral element is not necessary for the chiral induction, crucially both the ee and reaction rates can be increased without it.⁴ In the palladium-catalysed asymmetric hydrosilylation of functionalised styrenes, our ligands proved to be high yielding, and both regio- and enantioselective; phosphonite (S)-5a achieved ees of up to 95%. When the substrate used was 4-methoxystyrene, (S)-5a is significantly more enantioselective than Hayashi's benchmark phosphine (S)-H-MOP, demonstrating how MOP-phosphonite ligands have the potential to fill voids left by their phosphine counterparts in the field of asymmetric catalysis.

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Notes and references

- 1 *Phosphorus Ligands in Asymmetric Catalysis*, ed. A. Börner, Wiley-VCH, Weinheim, 2008; *Phosphorus(III) Ligands in Homogeneous Catalysis: Design and Synthesis*, ed. P. C. J. Kamer and P. W. N. M. van Leeuwen, John Wiley & Sons, Chichester, 2012.
- 2 T. Hayashi, *Acc. Chem. Res.*, 2000, **33**, 354; F. Lagasse and H. B. Kagan, *Chem. Pharm. Bull.*, 2000, **48**, 315; R. M. Hiney, A. Ficks, H. Müller-Bunz, D. G. Gilheany and L. J. Higham, in *Organometallic Chemistry*, ed. I. J. S. Fairlamb and J. M. Lynam, The Royal Society of Chemistry, Cambridge, 2011, vol. 37, pp. 27–45.
- 3 T. Hayashi, *Catal. Surv. Asia*, 1999, **3**, 127; J. W. Han and T. Hayashi, *Tetrahedron: Asymmetry*, 2014, **25**, 479.
- 4 A. Ficks, R. M. Hiney, R. W. Harrington, D. G. Gilheany and L. J. Higham, *Dalton Trans.*, 2012, **41**, 3515.
- 5 L. Bini, C. Muller and D. Vogt, *Chem. Commun.*, 2010, **46**, 8325.
- 6 J. I. van der Vlugt, J. M. J. Paulusse, E. J. Zipp, J. A. Tijmensen, A. M. Mills, A. L. Spek, C. Claver and D. Vogt, *Eur. J. Inorg. Chem.*, 2004, 4193.
- 7 C. Claver, E. Fernandez, A. Gillon, K. Heslop, D. J. Hyett, A. Martorell, A. G. Orpen and P. G. Pringle, *Chem. Commun.*, 2000, 961; M. T. Reetz and T. Sell, *Tetrahedron Lett.*, 2000, **41**, 6333.
- 8 R. M. Hiney, L. J. Higham, H. Müller-Bunz and D. G. Gilheany, *Angew. Chem., Int. Ed.*, 2006, **45**, 7248; A. Ficks, C. Sibbald, S. Ojo, R. W. Harrington, W. Clegg and L. J. Higham, *Synthesis*, 2013, 265.
- 9 P. G. A. Kumar, P. Dotta, R. Hermatschweiler, P. S. Pregosin, A. Albinati and S. Rizzato, *Organometallics*, 2005, **24**, 1306.
- 10 K. Kitayama, Y. Uozumi and T. Hayashi, *J. Chem. Soc., Chem. Commun.*, 1995, 1533.
- 11 H. L. Pedersen and M. Johannsen, *J. Org. Chem.*, 2002, **67**, 7982.
- 12 A. M. Fournier, R. A. Brown, W. Farnaby, H. Miyatake-Onzozabal and J. Clayden, *Org. Lett.*, 2010, **12**, 2222.
- 13 J. H. Hutchinson, T. J. Seiders, B. Wang, J. M. Arruda, J. R. Roppe and T. Parr, *Br. Pat.*, GB2470833, 2010.

