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## Structural ordering enhances highly selective production of acetic acid from CO<sub>2</sub> at ultra-low potential†

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Electrochemical reduction of CO<sub>2</sub> to value-added chemicals and fuels using renewable energy technologies is known to facilitate the creation of an artificial carbon cycle. Although the practical use of most conventional electrocatalysts is curbed by the low efficiency and poor stability of the catalyst there is also the need of large input energy in the form of potential. In this work, a family of bismuth-based transition metal chalcogenides was designed to enable multi-electron transfer for selectively reducing CO<sub>2</sub> to acetic acid at ultra-low potential of –0.1 V (vs. RHE). The structural design in AgBiS<sub>2</sub>, CuBiS<sub>2</sub> and AgBiSe<sub>2</sub> facilitated an optimized CO adsorption accounting for the production of acetic acid at low potential. The disordered arrangement of Ag and Bi in AgBiS<sub>2</sub> also favors CO hydrogenation, which leads to the formation of a large amount of methanol in addition to acetic acid. However, an induced structural ordering of these atoms upon selected substitution enhanced the lattice strain in CuBiS<sub>2</sub> and AgBiSe<sub>2</sub> favoring only C–C coupling and 100% acetic acid is produced at lower potential with stability up to 100 hours. The origin of the CO<sub>2</sub> reduced product has been validated by <sup>13</sup>CO<sub>2</sub> isotopic experiments and the mechanistic pathway has been proposed with the support of *in situ* IR experiments. Finally, a 4 times improvement in the current density of the best catalyst, AgBiSe<sub>2</sub>, was achieved in a flow cell configuration, which produced the highest ever acetic acid yield at lower potential with a faradaic efficiency of 49.81%. This work provides a novel strategy to improve electrochemical performance towards the formation of high value-added chemicals selectively at ultra-low potential.

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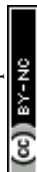
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### Broader context

This manuscript focuses on how structural ordering can facilitate the formation of acetic acid at ultra-low potential, which is a critical challenge in the electrochemical CO<sub>2</sub>RR. As per our knowledge, this is the first report of the production of acetic acid at ultra-low potential. Additionally, the manuscript provides an in-depth insight into the possible reaction mechanism using *in situ* IR along with deep understanding of how the crystal structures of ABiX<sub>2</sub> electrocatalysts tune the product selectivity and activity. This is the first work reporting long-term durability to produce acetic acid *via* eCO<sub>2</sub>RR. The origin of the CO<sub>2</sub> reduced product has been validated by <sup>13</sup>CO<sub>2</sub> isotopic experiments. Finally, the current density of the best catalyst has been improved by 4 times by performing the CO<sub>2</sub> reduction experiments in a flow cell configuration, which produced the highest ever acetic acid yield at lower potential with a faradaic efficiency of 49.81%. This work introduces a rational design of the catalyst to facilitate the C–C coupling reaction, and is expected to motivate researchers working in the area and can be a good guideline to rationally design and develop catalysts for a desired product from CO<sub>2</sub>.

The expedited technological advancement of modern society has led to severe environmental problems with global warming being the most pivotal amongst all others. Conversion of CO<sub>2</sub> into value-added chemicals and fuels appears to be one of the most attractive pathways to mitigate CO<sub>2</sub> accumulation in the atmosphere.<sup>1,2</sup> The electrochemical CO<sub>2</sub> reduction reaction (eCO<sub>2</sub>RR) facilitated by the usage of renewable energy resources is a sustainable alternative to large-scale dependence on fossil fuels and is advantageous since it can occur under ambient conditions and is solely dependent on potential bias.<sup>3–6</sup>

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Additionally, this strategy enables the on-site production of a wide variety of value-added chemicals and fuels, which includes C<sub>1</sub> and C<sub>2+</sub><sup>7</sup> hydrocarbons, mitigating their limitations in terms of distribution and storage specifically for toxic gases like CO.<sup>8</sup> However, the wide-scale application of the eCO<sub>2</sub>RR technology requires significant improvements in energy efficiency, catalyst stability and current density. The current most pivotal constraint lies with the high thermodynamic barrier required for the conversion of CO<sub>2</sub> to CO<sub>2</sub><sup>•−</sup> radicals, which is accompanied by the disruption of stable sp-hybridization to bend the linear CO<sub>2</sub> molecule.<sup>9</sup> Hence, CO<sub>2</sub> activation requires large activation energy barriers, which results in large overpotential requirements to drive the forward reaction. This makes it critical to design electrocatalysts that can facilitate stabilization of the intermediate to reduce the overpotential for the overall electrochemical process in addition to inhibiting the competitive hydrogen evolution reaction (HER).<sup>10</sup> Additionally the competition between the HER and eCO<sub>2</sub>RR stimulates low faradaic efficiencies (FEs) of certain products, which consumes most of the charge for this competitive reaction.<sup>11,12</sup> Therefore, it is desirable to develop electrocatalysts with low overpotential, fast kinetics and high efficiency for the conversion of CO<sub>2</sub> into multi-carbon products. Importantly, the production of C1 hydrocarbons<sup>13,14</sup> (CH<sub>4</sub>, HCOOH and CH<sub>3</sub>OH) is easier in comparison to C2 hydrocarbons<sup>15</sup> (C<sub>2</sub>H<sub>4</sub>, CH<sub>3</sub>COOH, C<sub>2</sub>H<sub>5</sub>OH), which have high volumetric energy densities and are the building blocks for the synthesis of long-chain hydrocarbon fuels.<sup>16,17</sup> The C2 product involves C–C bond formation which competes with the formation of C–H and C–O bonds under similar reaction conditions, which makes its production even more strenuous.<sup>18</sup> Hence, developing electrochemical systems that can mediate multi-proton transfer reactions at low overpotential still remains one of the most significant challenges. Amongst the C2 products, acetic acid is one of the most highly desired ones due to its immense applications in the food, pharma, chemical, textile, polymer, medicinal, and cosmetics sectors.<sup>19,20</sup>

However, as discussed earlier the eCO<sub>2</sub>RR to acetic acid involves a C–C coupling reaction along with adequate surface coverage of \*CO, \*CH<sub>2</sub> and \*CH<sub>3</sub> intermediates, which requires large overpotentials with much lower FE.<sup>21</sup> Fig. 1 and Table S1 (ESI†) summarize the FE for acetic acid as a function of potential, to date. Rational selection and design of the catalysts is very crucial for efficient and selective eCO<sub>2</sub>RR. Transition metal-based catalysts are often required for the eCO<sub>2</sub>RR in order to lower the activation energy barrier and drive the process at acceptable rates and at specific potentials.<sup>22</sup> Amongst the transition metals, Cu and its oxide derivatives have emerged as a unique class of electrocatalysts<sup>23–27</sup> that can efficiently convert CO<sub>2</sub> into a wide variety of hydrocarbons and their oxygenated products (CH<sub>3</sub>OH, CH<sub>4</sub>, C<sub>2</sub> and C<sub>2+</sub>).<sup>28,29</sup> Other transition metals like Au,<sup>30</sup> Ag,<sup>31</sup> and Zn<sup>32</sup> delivered high FE towards the formation of CO.<sup>33</sup> Additionally, P-block elements like Sn, Bi, In, Pb and their oxides have been known to catalyze the eCO<sub>2</sub>RR to produce formate or formic acid as the major product with FE as high as 90%.<sup>34–38</sup> The presence of lone pair



Fig. 1 Reported faradaic efficiency for the conversion of CO<sub>2</sub> to acetic acid as a function of applied potential on various catalysts.<sup>12,44–46</sup>

electrons in Bi<sup>3+</sup> accelerates the adsorption and activation of CO<sub>2</sub> molecules due to the Bi 6s and O 2p hybridization, which favors the formation of stereo chemically active lone pairs that promote electron donation to acidic adsorbed species such as CO<sub>2</sub>.<sup>39–41</sup> Chalcogens have been found to accelerate water activation. The promotional effect of sulfur on the indium surface to accelerate the reduction of CO<sub>2</sub> to formate has already been demonstrated by Ma *et al.*<sup>42</sup> The presence of negative charge on the chalcogen makes it nucleophilic and hence, traps CO<sub>2</sub>.<sup>43</sup> It is often wondered how these individual performances could be tuned selectively to a desired product upon a compound formed with a selected combination of these elements, which requires several controlled design strategies. In this work we tried to address three fundamental challenges in the eCO<sub>2</sub>RR: low CO<sub>2</sub> solubility, large overpotential and poor selectivity. To study these, we developed a series of MBiX<sub>2</sub> catalysts (M = Ag/Cu and X = S/Se). AgBiS<sub>2</sub> was chosen as the pristine catalyst for the eCO<sub>2</sub>RR and Ag was replaced with Cu to develop CuBiS<sub>2</sub> and S with Se to form AgBiSe<sub>2</sub>.

AgBiS<sub>2</sub>, CuBiS<sub>2</sub> and AgBiSe<sub>2</sub> were synthesized using a colloidal method with oleylamine as the solvent and reducing agent. PXRD patterns (Fig. S1, ESI†) demonstrated that AgBiS<sub>2</sub> crystallizes in the cubic phase (space group: *Fm3̄m*), CuBiS<sub>2</sub> in the orthorhombic phase (space group: *Pnma*) and AgBiSe<sub>2</sub> in the rhombohedral phase (*R3̄mh*). Fig. S2 (ESI†) shows the TEM images of AgBiS<sub>2</sub>, CuBiS<sub>2</sub> and AgBiSe<sub>2</sub>. The atomic coordinates and Wyckoff sites of all the catalysts have been tabulated in Table S2 (ESI†). As seen from Fig. S2a (ESI†), AgBiS<sub>2</sub> formed agglomerated spherical nanoparticles with an average particle size of 30–40 nm, while AgBiSe<sub>2</sub> showed the presence of irregular shaped nanostructures with an average size of 20–25 nm (Fig. S2c, ESI†). CuBiS<sub>2</sub> on the other hand formed a honey-comb like structure with a particle size of less than 10 nm. Selected area elemental mapping of the MBX<sub>2</sub> catalysts depicted uniform distribution of the respective elements *i.e.* M = Ag/Cu, Bi and X = S/Se throughout the nanoparticles (Fig. S3–S5, ESI†). EDAX taken on an ensemble of nanoparticles is in close agreement with these measurements confirming the expected stoichiometry of 1:1:2 in all three compounds



(Fig. S6, ESI<sup>†</sup>). The ABiX<sub>2</sub> catalysts were further characterized by XPS and XANES to understand their electronic structure (Fig. S7–S10 and Notes S1 and S2, ESI<sup>†</sup>).

The electrochemical CO<sub>2</sub>RR was carried out on the MBiX<sub>2</sub> catalysts in 0.5 M KHCO<sub>3</sub> with continuous CO<sub>2</sub> purging at 20 sccm flow rate (Fig. S11 and S12, ESI<sup>†</sup>). Fig. S13 (ESI<sup>†</sup>) depicts the observed current density at each applied potential on different MBiX<sub>2</sub> catalysts. At relatively lower overpotential (−0.1 V, −0.3 V vs. RHE), the current densities for all three chalcogenides are found to be similar. Chronoamperometry (CA) curves on different chalcogenide (Fig. S13b–d, ESI<sup>†</sup>) catalysts have been shown at different potentials in CO<sub>2</sub>-saturated 0.5 M KHCO<sub>3</sub> solution, indicating good electrochemical stability of our catalyst. Fig. S14 (ESI<sup>†</sup>) shows LSV polarization curves in which the increase in current density when the atmosphere

was changed from N<sub>2</sub> to CO<sub>2</sub> is apparent. The positive shift in onset potential upon saturating the electrolyte with CO<sub>2</sub> indicates the dominance of the CO<sub>2</sub> reduction process relative to the HER.

ECSA calculations demonstrate that AgBiSe<sub>2</sub> has the highest electrochemically active surface area (Fig. S15, ESI<sup>†</sup>). Fig. 2 shows the FEs and eCO<sub>2</sub>RR product distributions as a function of applied potential from −0.1 V to −1.1 V over all three catalysts, which were calculated combining GC, HPLC and NMR analyses (Fig. S16–S20, and Notes S3, ESI<sup>†</sup>). The product distribution and their corresponding FEs are given in Table S3 (ESI<sup>†</sup>). AgBiS<sub>2</sub> produced CH<sub>3</sub>OH at all the potentials with maximum FE (60.39%) at −0.3 V (Fig. 2a and Table S3, ESI<sup>†</sup>). On the other hand, acetic acid was formed at extremely low overpotentials (−0.1 V vs. RHE) with FE of 12.55%, which

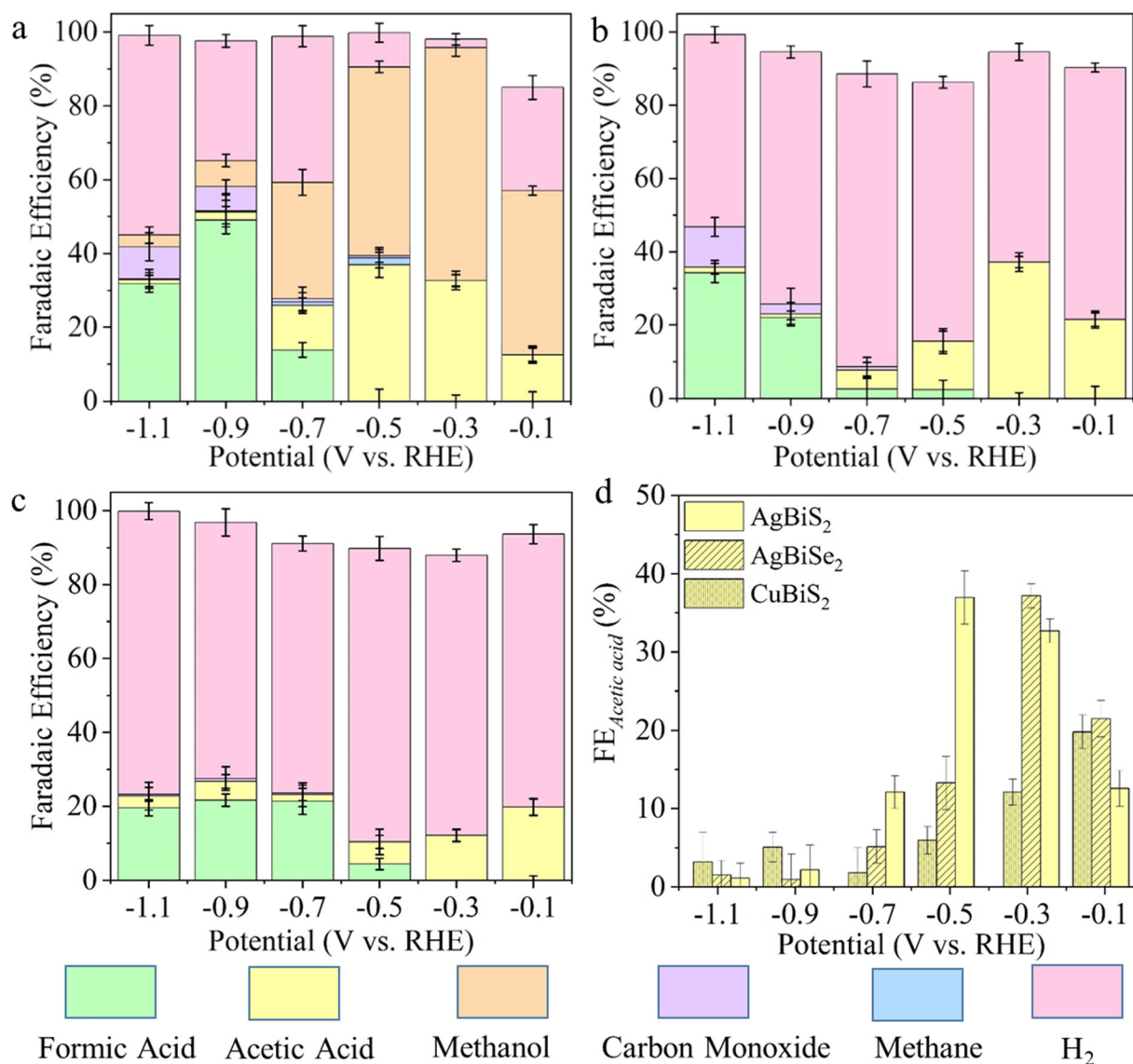


Fig. 2 Faradaic efficiency for each CO<sub>2</sub> reduced gaseous product (carbon monoxide, methane, hydrogen) and major liquid product (methanol, formic acid and acetic acid) as a function of potential after 1 hour of CA during the eCO<sub>2</sub>RR on (a) AgBiS<sub>2</sub>, (b) AgBiSe<sub>2</sub> and (c) CuBiS<sub>2</sub>, and (d) comparison of FE for acetic acid at different potentials for all the catalysts.



gradually increased to its maximum value of 35.97% at  $-0.3$  V.  $\text{AgBiSe}_2$  shows a similar trend of production of  $\text{CH}_3\text{COOH}$  (Fig. 2b) where the FE of  $\text{CH}_3\text{COOH}$  is higher at lower overpotential ( $-0.1$  V to  $-0.3$  V) and decreased at higher overpotential ( $-0.5$  V and  $-1.1$  V). Compared to  $\text{AgBiS}_2$ ,  $\text{AgBiSe}_2$  does not produce any detectable amount of  $\text{CH}_4$  or  $\text{CH}_3\text{OH}$  in the entire potential window. Only acetic acid was formed as the liquid product at a potential of  $-0.1$  V and  $-0.3$  V with 21.5% and 37.18% FE, respectively, which is higher than that observed in the case of  $\text{AgBiS}_2$ . The production of CO and HCOOH was observed as similar in both catalysts. Similarly,  $\text{CuBiS}_2$  also yields  $\text{CH}_3\text{COOH}$  as the only liquid product at an extremely low overpotential of  $-0.1$  V (vs. RHE), but the maximum FE observed was 28.26%, which is less than that of Ag-based systems (Fig. 2c). To further verify that the product was derived from  $\text{CO}_2$  reduction, an isotope labelled  $^{13}\text{CO}_2$  study was performed on the  $\text{AgBiS}_2$  catalyst since it produced both acetic acid and methanol.  $^1\text{H}$  NMR spectra (Fig. S21, ESI $^\dagger$ ) demonstrated the H signal due to  $^{13}\text{CH}_3$  groups on acetic acid and methanol. For both acetic acid and methanol, the H signal splits into two peaks due to coupling with  $^{13}\text{C}$  atoms. The isotope labelled  $^{13}\text{CO}_2$  study using HPLC on  $\text{AgBiS}_2$  (Fig. S22, ESI $^\dagger$ ) further confirmed that acetic acid was produced from  $\text{CO}_2$  and not from any other chemical. Fig. S23 (ESI $^\dagger$ ) shows the formation rate of different  $\text{CO}_2$  reduced products as a function of potential, which is in accordance with the observed FE trend. From this controlled design strategy of the catalysts, it is very clear that a disordered system favored methanol formation in addition to acetic acid but the ordered nature of the atoms in  $\text{CuBiS}_2$  and  $\text{AgBiSe}_2$  facilitated C–C coupling and helped in the selective production of acetic acid at lower potential.

Fig. 2d depicts the trend of FE towards  $\text{CH}_3\text{COOH}$  as a function of potential on each catalyst.  $\text{AgBiS}_2$  and  $\text{AgBiSe}_2$  have exhibited a volcano kind of trend where  $\text{CH}_3\text{COOH}$  is formed at lower overpotential ( $-0.1$  to  $-0.3$  V) with large FE. On the other hand, the  $\text{CH}_3\text{COOH}$  FE in the case of  $\text{CuBiS}_2$  linearly increased from higher potential to the maximum at lower potential ( $-0.1$  V). To scrutinize the durability of the  $\text{ABiX}_2$  catalysts we performed an endurance electrolysis experiment at  $-0.3$  V vs. RHE and monitored the liquid products generated using HPLC and NMR. Prolonged electrolysis of six hours showed negligible degradation in current density for  $\text{AgBiS}_2$ ,  $\text{AgBiSe}_2$  and  $\text{CuBiS}_2$  electrocatalysts indicating their high stability (Fig. S24, ESI $^\dagger$ ). Acetic acid remained as the major  $\text{C}_2$  product in the  $\text{AgBiSe}_2$ ,  $\text{CuBiS}_2$  and  $\text{AgBiS}_2$  electrocatalysts upon prolonged electrolysis of six hours along with formic acid while methanol formation was still observed for  $\text{AgBiS}_2$ . Additionally, long-term electrolysis led to the formation of other  $\text{C}_2$  products: ethanol and diethylene glycol for all three electrocatalysts (Fig. S25, ESI $^\dagger$ ).

Additionally, we also exploited long-term electrolysis up to 100 hours for our best active catalyst  $\text{AgBiSe}_2$ .  $\text{AgBiSe}_2$  demonstrated enhanced activity and performance under  $\text{eCO}_2\text{RR}$  conditions. The catalyst was found to be durable up to 100 hours of electrolysis with negligible degradation in electrocatalytic activity. In addition to being durable up to 100 hours it was observed that the faradaic efficiency (FE) for acetic acid was

consistent during this prolonged period with FE being 37.16% and 40.63% at the end of 1 hour and 100 hours, respectively (Fig. S26, ESI $^\dagger$ ). Post-electrochemical XRD at the end of 100 hours revealed no notable structural changes during prolonged electrolysis except that negligible  $\text{Ag}_2\text{O}$  was formed at this potential of  $-0.3$  V vs. RHE (Fig. S27, ESI $^\dagger$ ). The formation of ethanol and diethylene glycol was also observed upon long term electrolysis up to 100 hours (Fig. S28, ESI $^\dagger$ ). The FE of the obtained liquid products after 100 hours of electrolysis on the  $\text{AgBiSe}_2$  catalyst is shown in Fig. S29 (ESI $^\dagger$ ).

The reaction pathways and intermediates involved in the  $\text{eCO}_2\text{RR}$  over the three chalcogenides were probed by *in situ* attenuated total reflection (ATR-IR) spectroscopy (Fig. 3 and Table S4, ESI $^\dagger$ ).<sup>47</sup> Over  $\text{AgBiSe}_2$ , at  $-0.1$  V vs. RHE, a strong band appears at  $1411\text{ cm}^{-1}$  (Fig. 3a) indicating the symmetric stretch mode of the carbon bound  $^*\text{COO}^-$  intermediate.<sup>48</sup> (All the surface bound species will be prefixed \* from hereafter.) The intensity of the  $\text{COO}^-$  peak increases with time coinciding with the high FE of acetic acid. This is in line with previous reports in which  $^*\text{COO}^-$  serves as the intermediate for acetic acid by further proton coupled electron transfer processes and C–C coupling.<sup>49</sup> The slight hump at  $1566\text{ cm}^{-1}$  is assigned to the asymmetric stretch of  $\text{COO}^-$ . The weak peak appearing at  $1289\text{ cm}^{-1}$  belongs to the O–H deformation of the surface bound  $\text{COOH}$  intermediate. The intensity of this peak also rises with time indicating that it is also involved in acetic acid formation. Water consumption associated with the HER is indicated by the H–O–H bend at  $1621\text{ cm}^{-1}$ .<sup>50</sup> This peak may be overlapping with the C=O stretch of the  $^*\text{COOH}$  intermediate as seen from the slight hump at  $1660\text{ cm}^{-1}$ .<sup>48</sup> A weak band appearing at  $2055\text{ cm}^{-1}$  can be ascribed to linearly adsorbed CO.<sup>47</sup> The intensity of the peak is quite low as free CO does not form at this potential. The presence of  $^*\text{CO}$  at low applied potential only on  $\text{AgBiSe}_2$  is in agreement with its high FE for acetic acid at this potential. The formed  $^*\text{CO}$  may immediately turn to  $^*\text{HCO}$ , another important intermediate in the  $\text{CO}_2$  reduction pathway.

The minor dip at  $2340\text{ cm}^{-1}$  indicates the consumption of  $\text{CO}_2$  as evident from previous reports.<sup>51</sup> In  $\text{AgBiS}_2$ , the three vibrational modes of  $^*\text{COOH}$  are clearly seen: O–H deformation, C–O stretch and C=O stretch at  $1289$ ,  $1380$  and  $1612\text{ cm}^{-1}$ , respectively. The absence of the negative H–O–H bend is associated with the HER, which is in line with the high overall FE of the  $\text{AgBiS}_2$  (Fig. 3b) at low applied potentials.<sup>52</sup> In  $\text{CuBiS}_2$ , peaks associated with  $^*\text{COO}^-$  and  $^*\text{COOH}$  groups are overlapping at around  $1400\text{ cm}^{-1}$  (Fig. 3c). The relatively low intensity of the  $\text{COO}^-$  peaks as compared to  $\text{AgBiSe}_2$  and the absence of a hump near the H–O–H bending region for the  $^*\text{COOH}$  intermediate as in  $\text{AgBiS}_2$  is commensurate with the low overall FE of  $\text{CuBiS}_2$ . The IR spectra of  $\text{CuBiS}_2$  in the range of  $-0.1$  to  $-0.5$  V vs. RHE (Fig. S30–S32, ESI $^\dagger$ ) also show that the intensity of the  $\text{COO}^-$  and  $\text{COOH}$  peaks is quite low as compared to  $\text{AgBiS}_2$  and  $\text{AgBiSe}_2$ . The declining peak intensity of  $\text{COO}^-$  with applied potential (Fig. S33, ESI $^\dagger$ ) depicts that the production of formic acid at high potentials is *via* a separate mechanism and not through the carbon bound intermediates.





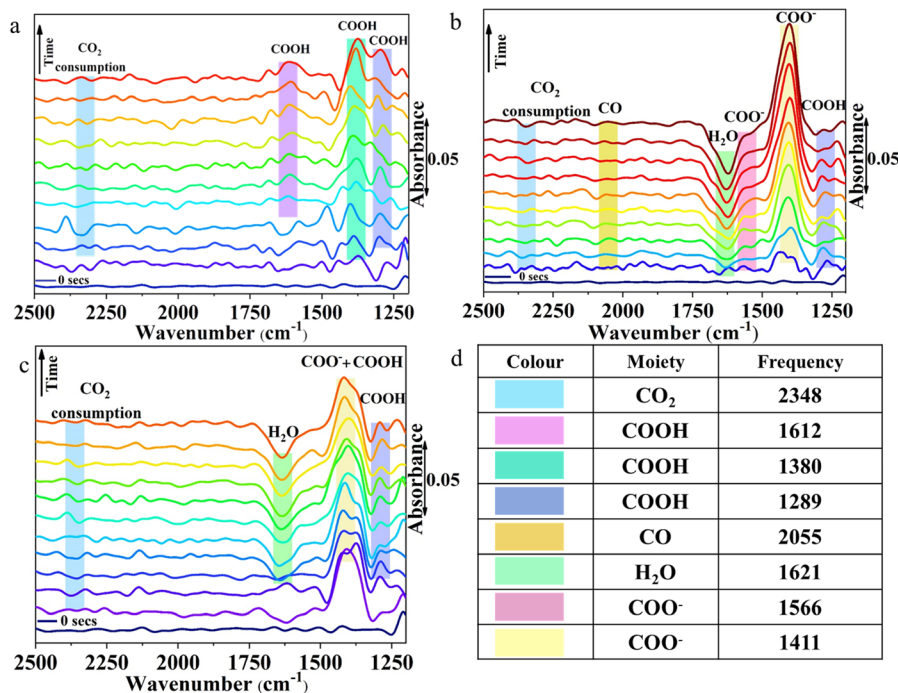


Fig. 3 (a–c) ATR-IR spectra during electrochemical CO<sub>2</sub> reduction on AgBiSe<sub>2</sub>, AgBiS<sub>2</sub>, and CuBiS<sub>2</sub> electrodes, respectively, at an applied potential of  $-0.1$  V vs. RHE and (d) adsorption frequencies of different intermediates adsorbed on the catalyst surface during CO<sub>2</sub> reduction.

To confirm that IR bands observed during *in situ* ATR arise due to CO<sub>2</sub> reduction and not from bicarbonate species, we have performed the ATR studies using 0.1 M KCl (Fig. S34, ESI<sup>†</sup>). Since, the IR bands appear almost at the same stretching frequencies as those in 0.5 M KHCO<sub>3</sub>, we can confirm that CO<sub>2</sub> reduction results in the formation of the observed \*COO<sup>-</sup> and \*COOH intermediates. Based on the results of the *in situ* ATR studies and with the aid of the previous reports, the mechanism for the conversion of CO<sub>2</sub> to acetic acid and methanol formation over the chalcogenide catalysts may be proposed (Fig. 4). The first step involved in CO<sub>2</sub> reduction is the activation of the CO<sub>2</sub> molecule. Here, CO<sub>2</sub> is adsorbed into the chalcogenide surface by accepting an electron as a carbon bound CO<sub>2</sub>\*<sup>-</sup> radical ion.<sup>53</sup> The LUMO of the activated CO<sub>2</sub> molecule is localized at C while the HOMO is localized at O due to the lower electronegativity of C in comparison to that of O. Hence, CO<sub>2</sub> is liable to undergo both nucleophilic and electrophilic reactions, respectively.<sup>43</sup> The adsorbed CO<sub>2</sub> molecule then undergoes hydrogenation to form a \*COOH moiety followed by proton coupled electron transfer processes that result in \*CHO, thereby converting to acetic acid and methanol in the ensuing C–C coupling process at lower potential.<sup>49</sup>

Based on earlier reports of chalcogenides with p-block elements, it can be predicted that the formation of HCOO\* on the S–Bi surface is expected to be exergonic while the formation of H\* from H<sub>2</sub>O is endergonic, thereby facilitating the formation of formic acid at higher potential.<sup>24</sup> To validate the mechanism further, we explored the crystal structure of all three catalysts. As seen from Fig. 5, AgBiSe<sub>2</sub> and CuBiS<sub>2</sub> have ordered structures with shortest X–Bi (X = S/Se) bond distances

of 3.04 Å and 3.15 Å, respectively. The shorter bond distance aids in the close proximity of adsorbed C atoms on AgBiSe<sub>2</sub> and CuBiS<sub>2</sub> surfaces, thereby favoring only C–C coupling at lower potential and accounting for its high selectivity towards acetic acid as compared to AgBiS<sub>2</sub>. Also, the adsorption energy of reaction intermediates plays a pivotal role in product selectivity. A lower d-band center indicates stronger adsorption of the reaction intermediate and weak desorption ability.<sup>22</sup> Since C–C coupling is a necessity for the formation of C<sub>2</sub> products, stronger CO adsorption will facilitate the coupling. This makes it important to analyze the local charge distribution in the lattice. To have a qualitative understanding of the variation in localized charge distribution of the three different lattice systems, XPS and XANES analyses were performed (Fig. S7–S10 and Notes S1 and S2, ESI<sup>†</sup>). Fig. S35 (ESI<sup>†</sup>) shows the valence band spectra of AgBiS<sub>2</sub>, AgBiSe<sub>2</sub> and CuBiS<sub>2</sub> derived from XPS, which shows that AgBiSe<sub>2</sub> has the lowest d-band center followed by CuBiS<sub>2</sub> and it is high in the case of AgBiS<sub>2</sub>. CO adsorption is expected to be the strongest in AgBiSe<sub>2</sub> with weak desorption. Hence, C–C coupling in AgBiSe<sub>2</sub> is more favored accounting for its high selectivity followed by CuBiS<sub>2</sub>. However, since AgBiS<sub>2</sub> has the highest bond distance, CO hydrogenation will be kinetically more favored as compared to C–C coupling. CO hydrogenation results in the formation of \*CHO or \*COH intermediates resulting in the formation of methanol and methane. This rationale validates our experimental findings of AgBiS<sub>2</sub> favoring the formation of methanol and methane in addition to acetic acid. XANES and XPS analyses further shed light on the activity difference between the ordered structures AgBiSe<sub>2</sub> and CuBiS<sub>2</sub>. Fig. 5d and e shows



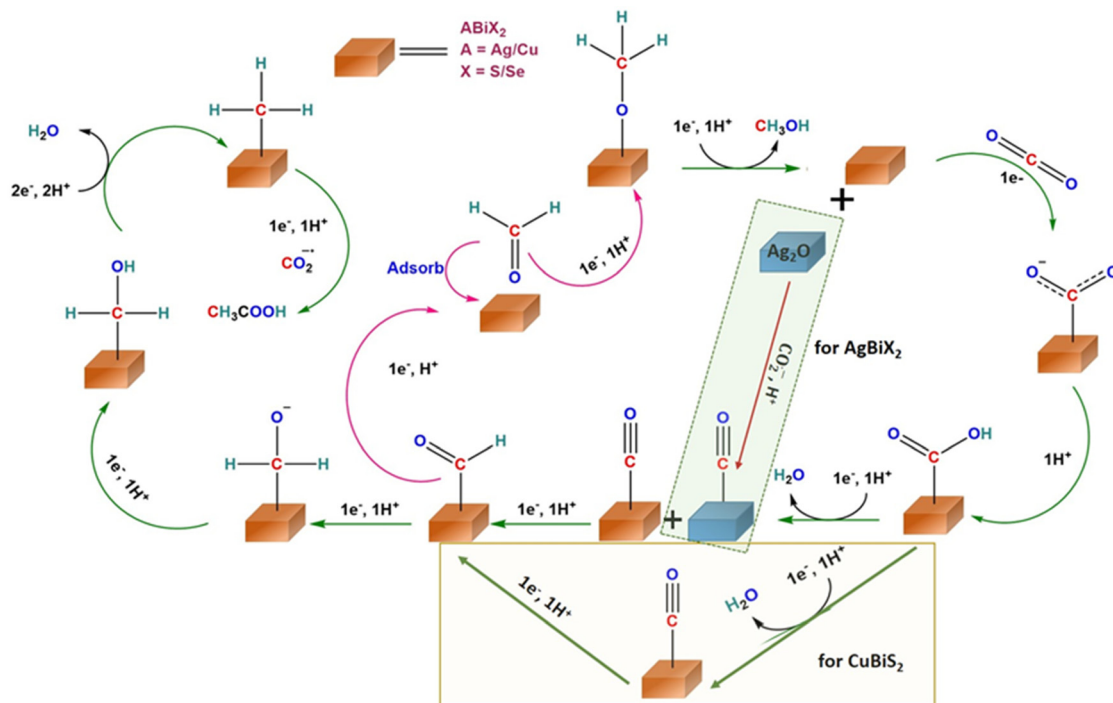


Fig. 4 Proposed reaction mechanism of the eCO<sub>2</sub>RR for the production of CH<sub>3</sub>COOH and HCOOH on CuBiS<sub>2</sub>, AgBiSe<sub>2</sub>, and AgBiS<sub>2</sub> predicted from *in situ* IR experiments.

a downshift in both binding energy and absorption edge of Bi in AgBiSe<sub>2</sub> relative to that of CuBiS<sub>2</sub>. This indicates that Bi in AgBiSe<sub>2</sub> has an oxidation state of 3<sup>δ-</sup> instead of 3<sup>+</sup> which makes Se more electronegative. On the contrary Cu in CuBiS<sub>2</sub> is

expected to have an oxidation state in between 0 and +1 due to the presence of Bi in the 3<sup>δ+</sup> state.

A careful investigation of the crystal structure (Fig. 5b) demonstrates that no bond exists between Ag and Bi in AgBiSe<sub>2</sub>

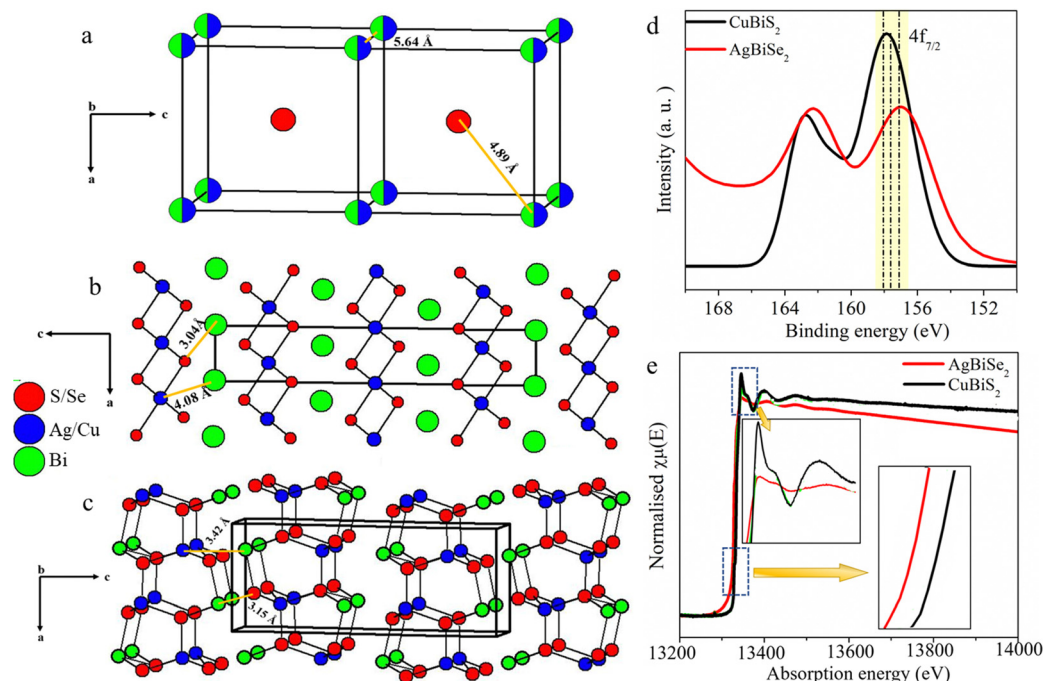


Fig. 5 Crystal structures of (a) AgBiS<sub>2</sub>, (b) AgBiSe<sub>2</sub> and (c) CuBiS<sub>2</sub>. Comparison of XPS spectra of Bi 4f (d) and XANES spectra of the Bi-L III edge (e) for AgBiSe<sub>2</sub> and CuBiS<sub>2</sub> electrocatalysts.



while a Cu–Bi bond with bond distance of 3.4 Å exists in CuBiS<sub>2</sub>. Hence, correlating the electronic and crystal structure of both the ordered compounds, *i.e.* AgBiSe<sub>2</sub> and CuBiS<sub>2</sub>, we anticipate that facile charge transfer between Bi and Se in AgBiSe<sub>2</sub> expedites easy adsorption of the C-bound \*COOH intermediate, which rationalizes the higher FE for acetic acid in AgBiSe<sub>2</sub> relative to that of CuBiS<sub>2</sub>.<sup>54,55</sup>

Since structural changes can happen at the reducing voltage in the CO<sub>2</sub> reduction reaction, powder XRD patterns of the working electrodes were recorded immediately after the electrolysis at each potential to understand the active crystallographic phase during that potential. During the electrochemical measurements on AgBiX<sub>2</sub> compounds, the formation of Ag<sub>2</sub>O was observed (Fig. S36a and b, ESI<sup>†</sup>) whereas CuBiS<sub>2</sub> did not undergo any structural change (Fig. S36c, ESI<sup>†</sup>). AgBiSe<sub>2</sub> has not changed its structure at any potential, but AgBiS<sub>2</sub> became structurally ordered upon increasing the applied potential.

This strongly confirms the earlier explanation that AgBiS<sub>2</sub> favors the production of both acetic acid and methanol at lower potential because of the presence of both ordered and disordered phases. Additionally, in both cases the evolution of a small amount of Ag<sub>2</sub>O facilitates CO<sub>2</sub> to CO conversion, which is an intermediate step in acetic acid formation. Because of this additional reaction, we have higher FE towards acetic acid in the case of Ag compounds compared to Cu compounds. Post electrolysis XPS analysis was done to understand the changes in electronic structure under CO<sub>2</sub>RR electrolysis conditions (Fig. S37 and S38, ESI<sup>†</sup>). As evident from Fig. S39 (ESI<sup>†</sup>), Ag surfaces in ABiX<sub>2</sub> catalysts undergo a shift to lower binding energies upon application of negative potential, which is a good indication of significant reduction of metallic surfaces during the eCO<sub>2</sub>RR.<sup>56</sup> Besides selectivity and activity, stability is another crucial descriptor while evaluating the performance of electrocatalysts for the eCO<sub>2</sub>RR.

To minimize CO<sub>2</sub> mass transport issues and achieve high current density we explored the CO<sub>2</sub>RR activity of the MBiX<sub>2</sub> catalysts in a flow-cell configuration (Fig. S40a–c, ESI<sup>†</sup>) as compared to the H-cell configuration. Flow cells with a gas diffusion electrode (GDE) can enable efficient transport of CO<sub>2</sub> to the gas–electrolyte–electrode interface and minimize the diffusion layer thickness to the nanometer scale.<sup>57,58</sup> A flow-cell configuration helped in achieving current densities which are almost 4-times as compared to that in the H-cell<sup>59,60</sup> (Fig. S41 and Table S5, ESI<sup>†</sup>). The flow-cell led to the formation of acetic acid and formic acid as the major liquid products as was observed from the H-cell. However, interestingly, product analysis showed a considerable reduction in H<sub>2</sub> FE while the FE of acetic acid was improved significantly (Fig. 6). Our best catalyst, AgBiSe<sub>2</sub> demonstrated acetic acid as the CO<sub>2</sub> reduced product with FE as high as 49.81% at an ultra-low potential of –0.3 V (Fig. S6b, ESI<sup>†</sup>). The achieved FE for acetic acid from AgBiSe<sub>2</sub> in the flow-cell was almost 4 times that achieved from the H-cell at a current density of 15.8 mA cm<sup>–2</sup>. The increased selectivity to acetic acid and decreased selectivity to hydrogen is likely due to the higher local pH in the flow cell

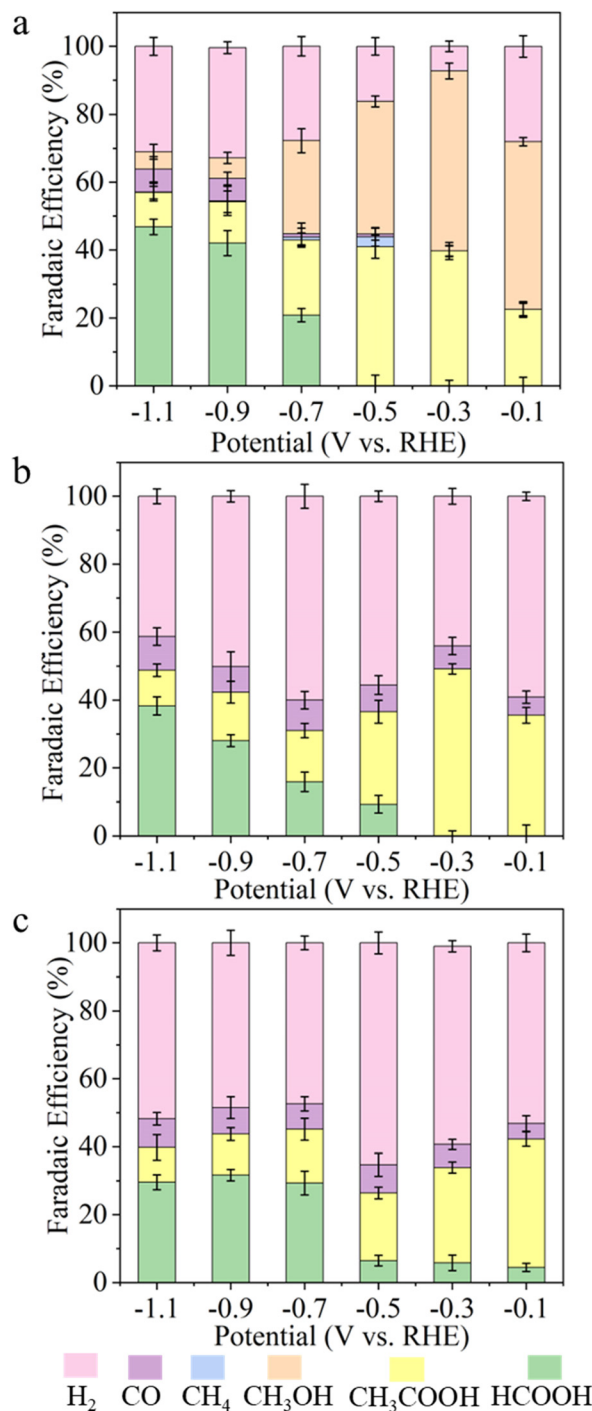


Fig. 6 Faradaic efficiency for all the liquid and gaseous products found during CO<sub>2</sub> reduction as a function of potential on (a) AgBiS<sub>2</sub> and (b) AgBiSe<sub>2</sub> and (c) CuBiS<sub>2</sub> catalysts in the flow cell configuration in 0.5 M KHCO<sub>3</sub>.

microenvironment.<sup>61</sup> This accelerates the OH-mediated nucleophilic attack of the ethenone intermediate leading to acetic acid.<sup>62</sup> The maximum achieved FE for acetic acid at a current density of 15.8 mA cm<sup>–2</sup> was found to be stable up to 40 hours in the flow-cell configuration (Fig. S42, ESI<sup>†</sup>).



In conclusion, we have rationally designed a set of catalysts to tune the selective production of acetic acid from CO<sub>2</sub> at lower potential. Our controlled studies clearly manifest that atomic ordering and optimized chemical bonding are very crucial parameters in controlling the reaction pathway to a desired product. We have chosen a couple of the most studied transition metals for CO<sub>2</sub> reduction (Cu and Ag) and alloyed them with Bi and chalcogens, which helped in tuning the global and local structure. The atomically ordered AgBiS<sub>2</sub> and AgBiSe<sub>2</sub> compounds favored the selective production of acetic acid with significant FE at low overpotential. *In situ* IR studies mapped the reaction pathways with each step and important intermediates for the formation of products at various potentials.

## Conflicts of interest

There are no conflicts to declare.

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