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Two-dimensional materials as catalysts, interfaces, and electrodes for an efficient hydrogen evolution reaction

Yun Seong Cho and Joohoon Kang *

Two-dimensional (2D) materials have been significantly investigated as electrocatalysts for the hydrogen evolution reaction (HER) over the past few decades due to their excellent electrocatalytic properties and their structural uniqueness including the atomically thin structure and abundant active sites. Recently, 2D materials with various electronic properties have not only been used as active catalytic materials, but also employed in other components of the HER electrodes including a conductive electrode layer and an interfacial layer to maximize the HER efficiency or utilized as templates for catalytic nanostructure growth. This review provides the recent progress and future perspectives of 2D materials as key components in electrocatalytic systems with an emphasis on the HER applications. We categorized the use of 2D materials into three types: a catalytic layer, an electrode for catalyst support, and an interlayer for enhancing charge transfer between the catalytic layer and the electrode. We first introduce various scalable synthesis methods of electrocatalytic-grade 2D materials, and we discuss the role of 2D materials as HER catalysts, an interface for efficient charge transfer, and an electrode and/or a growth template of nanostructured noble metals.

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Introduction

Electrochemical water splitting is an environmentally friendly way for green hydrogen production through the hydrogen evolution reaction (HER).¹ To enable the reaction, the required thermodynamic Gibbs free energy (ΔG) is approximately 237.2 kJ mol⁻¹, corresponding to the uphill reaction with a standard potential (ΔE) of 1.23 eV compared to the reversible hydrogen electrode (RHE).² The HER includes the Volmer step, the Heyrovsky step, and the Tafel step for each acidic and alkaline condition, expressed as follows:³

School of Advanced Materials Science and Engineering, Sungkyunkwan University (SKKU), Suwon 16419, Republic of Korea. E-mail: joohoon@skku.edu



Yun Seong Cho

Yun Seong Cho is currently a Ph.D. student under the supervision of Prof. Joohoon Kang at the School of Advanced Materials Science & Engineering in Sungkyunkwan University (SKKU). He received his B.S. degree from the School of Advanced Materials Science & Engineering in 2021 from SKKU. His research interests include the synthesis, processing, and electrochemical applications of nanomaterials.



Joohoon Kang

Prof. Joohoon Kang is an Assistant Professor in the School of Advanced Materials Science and Engineering and the KIST-SKKU Carbon Neutral Research Center at Sungkyunkwan University (SKKU). He received his B.S. and M.S. degrees in materials science and engineering from Yonsei University in Korea in 2009 and 2011, respectively, and his Ph.D. degree in materials science and engineering from Northwestern University in 2018. He then moved to the University of California at Berkeley as a postdoctoral fellow in the College of Chemistry. His research interests include the synthesis, processing, and (opto) electronic applications of nanomaterials.

In acidic electrolytes:

Volmer step



Heyrovsky step



Tafel step



In alkaline electrolytes:

Volmer step



Heyrovsky step



Tafel step



For efficient hydrogen production based on the above reactions, many considerable factors, including the kinetics of each step, electron mobility, adsorption and desorption of hydrogen ions, and the separation rate of generated hydrogen gas, can contribute to increase ΔE higher than 1.23 eV as an activation barrier of the reaction. Therefore, it is highly necessary to investigate a comprehensive system for HER (*i.e.*, catalytic electrode), consisting of not only catalytic materials, but also interfacial layers and conducting electrodes, to fully optimize such factors and the resulting HER efficiency.

A straightforward approach to develop catalytic electrodes is combining different materials for exploiting each desired property. This strategy is synergetic in the way of integrating a conductive electrode component that facilitates efficient charge transfer with catalytic materials triggering active reactions at the electrolyte interface. For such synergetic catalytic electrodes' preparation, various carbon supports have been used to ensure efficient electron flow as an electrode part.^{4,5} In the catalyst part, where the reaction occurs directly upon contact with the electrolyte, achieving high activity is important, with predictions primarily based on the hydrogen adsorption free energy.⁶ The well-known catalysts with superior electrocatalytic performances (*i.e.*, a low overpotential and small Tafel slope) are generally based on noble metals (*e.g.*, platinum, Pt).⁷ Regardless of the remarkable performances, their high cost and scarcity in quantity highly motivate us to minimize their use by maximizing effective active sites (*i.e.*, single-atom catalysts) or discover other cost-effective material candidates to replace noble metals.^{8,9} Lastly, a new perspective is proposed to enhance the performance of existing catalysts by designing a functional interfacial layer between electrodes and catalysts to control over the desired charge transfer. The flow of electrons enhanced through semiconductor heterojunction structures is a notable example and has been actively studied in recent years.^{10,11}

For more than a decade, two-dimensional (2D) materials have received significant attention as an alternative materials family to replace the conventional noble metal-based catalysts due to their promising electrocatalytic properties, such as a large surface area and ease of charge transport.^{12,13} The atomically thin and unique structure ensures that even as a monolayer, all edge active sites are exposed, resulting in significant catalytic efficiency.¹⁴ In particular, 2D structures minimize dangling bonds, leading to reduced electron scattering and improved electrical properties unlike other dimensional materials.^{15,16} These properties correspond to the catalytic role, which requires effective charge transfer to the active sites. Moreover, the ease of sheet-to-sheet electrical connectivity through van der Waals (vdW) interactions and the scalability to wafer-scale make 2D structures a promising platform for many electronic applications.¹⁷ As a low-energy approach to material integration, vdW integration represents the superiority of 2D materials in terms of cost-effectiveness.^{18,19} This physical assembly method does not rely on one-to-one chemical bonding and does not involve direct chemical processing of existing materials. Furthermore, electrocatalytic-grade 2D materials can be easily synthesized in a large quantity at a relatively low cost compared to noble metals. Toward this end, solution-based processing of 2D materials such as liquid-phase exfoliation (LPE), alkali metal intercalation, and molecular intercalation-based electrochemical exfoliation may be one of the most suitable approaches for scalable synthesis of 2D materials for electrocatalytic applications.^{20,21}

A broad range of electronic properties of 2D materials including metallic, semiconducting, and insulating properties are desirable to implement in each part of the catalytic electrode, which is mainly composed of catalysts, interfaces, and electrodes.²² The use of 2D electrode materials, such as graphene, allows the fabrication of catalytic electrodes with thickness advantages, which demonstrate excellent HER catalytic properties when combined with noble metals.^{23,24} Additionally in the catalyst part, strategies such as defect engineering have been employed to maximize the active sites of 2D catalytic materials and enhance the HER activity, leading to significant performance improvements.^{25,26} A semiconductor heterojunction structure, embodied as a p-n junction, facilitates efficient charge transfer and functionalization, thanks to its internal built-in potential.¹⁸ While the synthesis of heterostructures such as p-n junctions for HER catalytic electrodes often involves complex processes of chemical integration into the base material, vdW heterostructures can be achieved simply by vertically stacking 2D sheets.^{27,28} vdW heterostructures are used in various electronic applications due to their excellent electrical conductivity, suggesting that they will be actively applied as catalytic electrodes in the future.²⁹ However, it has not yet been fully elucidated to use 2D materials as an interfacial layer due to the lack of fully optimized scalable synthesis methods.

Here, we provide a thorough review on the roles of 2D materials with diverse electronic properties in improving the HER efficiency as catalytic materials, interlayers, and conduc-

tive electrodes, respectively. For example, 2D semiconducting transition metal dichalcogenides (TMDCs) such as molybdenum disulfide (MoS_2) are the most intensively studied 2D catalytic materials.³⁰ In particular, several synthesis methods of 2D TMDCs for HER have been reported to maximize active sites (e.g., edges or defects on a basal plane) as well as to induce phase transformation from the semiconducting 2H phase to the metallic 1T phase for efficient charge transport.^{31,32} Besides, 2D semiconducting materials with different electronic types can be simultaneously placed at the electrode/catalyst interface by forming a p-n junction.²⁸ Finally, 2D metallic materials have been used as conductive electrodes or templates for the synthesis of other catalytic materials with controlled nanoscale morphologies.³³

2D material-based catalytic electrodes for HER

A catalytic electrode for HER can be mainly categorized into three parts as illustrated in Fig. 1a: (i) a catalyst layer that directly interfaces with the electrolyte; (ii) an interlayer where the charge transfer occurs; (iii) a conductive electrode layer which primarily supplies electrons.^{34,35} With a rapid emer-

gence of 2D materials research, various 2D materials with desired electronic properties can be directly implemented to each layer for improving HER properties.

TMDCs have been the most widely studied catalytic 2D materials family for HER because of the expected free energy of hydrogen adsorption on the surface, which is a key parameter in HER kinetics.^{32,36–38} As shown in Fig. 1b, the relationship between the HER efficiency and the hydrogen adsorption energy was quantified as a ‘volcano plot’, which is a function of the catalyst’s exchange current density and hydrogen adsorption free energy (ΔG_{H^*}).¹² When $\Delta G_{\text{H}^*} < 0$, the rates of both Tafel and Heyrovsky reactions, which involve the combination of hydrogen adatoms into hydrogen molecules, can be constrained. Conversely, when $\Delta G_{\text{H}^*} > 0$, the rate of the Volmer reaction, which involves the formation of hydrogen adatoms, is slowed down. As a result, catalysts placed at $\Delta G_{\text{H}^*} \approx 0$ in the volcano plot, where the hydrogen adatom is moderately bound, are considered as strong candidates for HER, in accordance with the Sabatier principle.³⁹ This principle underscores the importance of maintaining a delicate balance in catalyst–reactant interactions, as it influences the catalyst’s effectiveness in the reaction.³⁹ In this context, TMDCs such as 1T- WS_2 , MoSe_2 , and MoS_2 are considered as starting materials for HER catalysts, and further modifications such as size selec-

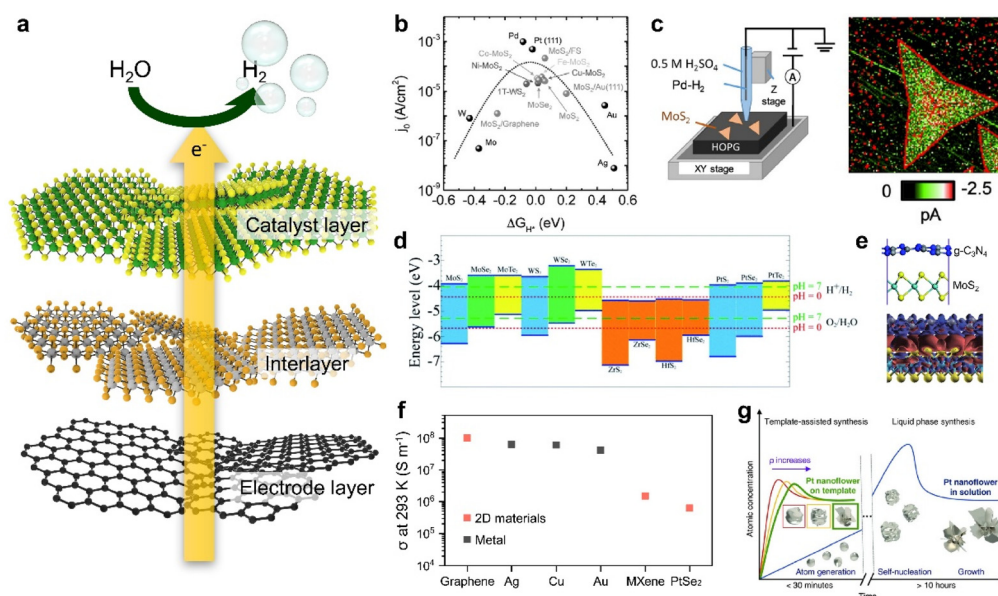


Fig. 1 (a) Schematic illustration of the electrocatalysis electrode for the HER consisting of a catalyst layer, an interlayer, and an electrode layer. (b) Volcano plot illustrating the relationship between the exchange current density and the free energy of hydrogen adsorption for different catalysts: free-standing MoS_2 edges, MoS_2 edges on Au (111), MoS_2 edges on graphene, vertically grown MoS_2 edges, vertically grown MoS_2 edges doped with Fe, Cu, Ni and Co, vertically grown MoSe_2 edges, and a 1T WS_2 basal plane. Various metal catalysts are also presented for comparison. Data sources: ref. 40–42. Reproduced with permission from ref. 12 Copyright 2016 John Wiley and Sons. (c) Scanning electrochemical cell microscopy (SECCM) set up and the corresponding SECCM current image of MoS_2 nanosheets during the HER. Reproduced with permission from ref. 44 Copyright 2020 John Wiley and Sons. (d) Band edge positions of single-layer TMDCs relative to the vacuum level. The redox potentials of water splitting at pH 0 (red dotted line) and pH 7 (green dashed line) are also shown for comparison. Reproduced with permission from ref. 49 Copyright 2013 American Chemical Society. (e) Scheme describing the $\text{g-C}_3\text{N}_4/\text{MoS}_2$ nanocomposite and charge density differences in the $\text{g-C}_3\text{N}_4/\text{MoS}_2$ nanocomposite. The red and blue regions represent charge accumulation and depletion, respectively, for the isosurface value set at $0.001 \text{ e } \text{\AA}^{-3}$. Reproduced with permission from ref. 52 Copyright 2013 Royal Society of Chemistry. (f) Electrical conductivities of graphene, Ag, Cu, Au, MXenes, and PtSe_2 . Data sources: ref. 59–63. (g) La Mer diagrams describing the mechanism of nucleation and subsequent growth of Pt nanoflowers for different synthesis methods. Reproduced with permission from ref. 33 Copyright 2022 The Authors.

tion, defect formation, and doping have been attempted to achieve better catalytic activity.^{40–43} Such efforts to maximize catalytic active sites in 2D nanosheets can be resolved by investigating high-resolution scanning electrochemical cell microscopy (SECCM) analysis (Fig. 1c).⁴⁴ This analysis clearly visualizes that the active sites are structurally concentrated at the edge sites of 2D TMDCs by using high-resolution SECCM imaging.^{44,45}

A broad range of semiconducting 2D materials also can be potentially utilized as an interlayer between a catalytic layer and an electrode to improve charge transfer. From energy band alignment-based simple expectations, specific 2D materials with desired electronic band structures can be easily chosen as an interlayer from the entire family with diverse electronic properties with respect to their elemental compositions.^{46,47} The energy band structure of 2D materials can be further modulated by controlling the number of layers (*i.e.*, thickness) in their atomically thin limit due to quantum confinement effects.^{47,48} As shown in Fig. 1d, a list of TMDCs possessing diverse energy bandgap properties can be placed at the interface between a catalytic layer and an electrode with a vertically-stacked catalytic electrode configuration to boost up charge transfer.⁴⁹ For example, the HER performance and stability can be highly improved by designing a facile transfer of electrons from a high conduction band to stepwise lower conduction bands with an energy level where the HER occurs.^{50,51} Furthermore, this approach enables us to obtain higher energy efficiency beyond the applied voltage by directly utilizing sunlight to transfer excited charges to a catalyst surface through an interlayer (Fig. 1e).^{52,53}

Lastly, metallic 2D materials such as graphene and MXenes have been investigated as conductive electrode layers.^{54–56} In addition to their intrinsically high electrical conductivity, their assembled thin film structure with vdW interfaces has the structural advantage of minimizing the charge momentum loss at the nanosheet-to-nanosheet interfaces and providing very clean interfaces that promote charge transport, preserving high electrical conductivity (Fig. 1f).^{57,58} Graphene, in particular, exhibits the electrical conductivity surpassing that of metals like Ag, Cu, and Au as commonly used electrodes.^{59–61} Other 2D materials like MXenes and PtSe₂ have similarly shown high electrical conductivity, as the emergence of innovative electrodes.^{62,63} Besides, such conductive 2D materials can be adopted as templates for synthesizing metallic nanoparticles, aiming to control their shapes by leveraging distinct electrical characteristics (Fig. 1g).³³ The reaction rate, a key variable to determine final nanostructures, can be easily controlled by the electrical resistance of the template.³³

2D materials as catalyst layers

Synthesis methods of 2D materials

With several decades of history of 2D materials research, various synthesis methods have been reported as illustrated in Fig. 2. As described in the *x*- and *y*-axis, respectively, the pro-

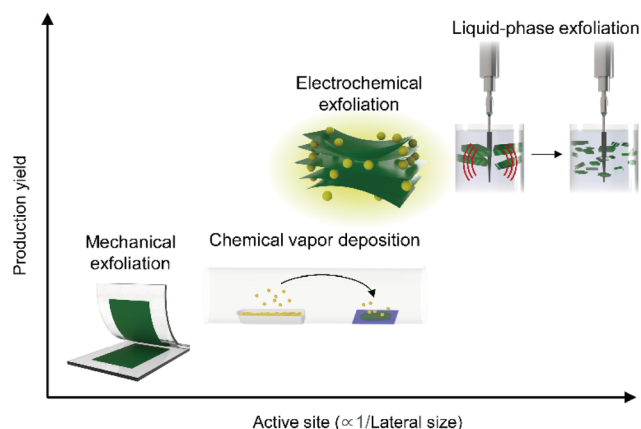


Fig. 2 Plot comparing the production yields and relative catalytic active site densities of 2D materials produced by various synthesis methods.

duction yield and active site density for HER (inversely proportional to the lateral size, in general) have a trade-off relationship. Since the first micromechanical exfoliation of monolayer graphene using Scotch tape by Novoselov *et al.* in 2004,⁶⁴ many fundamental studies including electrochemical reactions from individual nanosheets have been pursued to achieve a systematic understanding of the underlying mechanisms of hydrogen evolution.^{65,66} However, this synthesis method limits to utilize 2D materials in practical applications (*i.e.*, industrial level production of hydrogen gas) due to the lack of scalability. To overcome this limitation, a scalable synthesis method of 2D materials *via* chemical vapor deposition (CVD) processes has been proposed, enabling the synthesis of 2D materials in wafer-scale.⁶⁷ However, the CVD method is strongly dependent on the growth substrate, and thus this approach further requires to transfer to evaluate the HER performance of 2D materials. Considering the optimal samples for HER applications, both the abovementioned approaches are not fully satisfying these requirements: (i) the approach for synthesizing smaller lateral sizes to maximize active site densities has been favoured;^{45,68,69} (ii) given the necessity for competitive pricing to replace costly noble metal catalysts, the production yield, indicating the process efficiency, has been selected as a key metric;⁷⁰ (iii) specifically, for TMDCs with active sites primarily located at the edges or basal plane defects, scalable synthesis processes inducing smaller lateral sizes and vacancies might be more advantageous.⁷¹ From this perspective, solution-based processing stands out as a suitable candidate such as LPE, alkali metal intercalation, and molecular intercalation-based electrochemical exfoliation.⁷² These solution-based processing methods enable the production of electrocatalytic-grade 2D materials in a large quantity as a stable dispersion form in a cost-effective way. LPE is used to produce 2D materials from their bulk crystal or powder by application of ultrasonication, shear mixing, and ball milling. The resulting materials possess relatively smaller lateral dimensions (<100 nm) with a broad range of thickness distributions.²¹ 2D crystals can be also exfoliated by intercalation of

alkali metals (*e.g.*, Li, Na, and K) and stabilized in a solvent.⁷³ In this case, the resulting 2D semiconductors possess relatively large lateral sizes ($\sim 1\ \mu\text{m}$) in their atomically thin limit, but interestingly exhibit metallic behaviour in their electronic properties originated from the phase transformation (*i.e.*, 2H to 1T). While the alkali metal intercalation-driven 2D materials exfoliation results in a metallic behaviour of the samples, another type of intercalation method, molecular intercalation-based electrochemical exfoliation, can produce 2D samples with similar structural characteristics to the alkali-metal intercalated ones (*i.e.*, relatively large lateral sizes ($\sim 1\ \mu\text{m}$) in their atomically thin limit), but exhibiting intrinsic semiconducting properties without phase transformation.²⁰ The abovementioned solution-based processing methods for 2D materials have been commonly explored for scalable production of 2D materials for HER applications. However, the distinctive properties in the structure and electrical behaviour of the resulting materials have been considered for their use in each part of the catalytic electrode to maximize their potential for HER performance.

Demonstration and practical application of catalytic 2D materials

For decades, MoS_2 has been attracting significant attention as a catalytic 2D semiconductor for HER due to its promising catalytic properties and cost-effectiveness.⁷⁴ As the edge sites of MoS_2 have been expected to be the active site of the electrochemical reaction, while the basal plane has been considered as electrochemically inert,⁴⁵ it was first demonstrated through a microcell fabricated through selected-area lithography (Fig. 3a).¹⁴ To compare the HER performance of each basal plane and the edge sites of the CVD-grown monolayer MoS_2 , all samples except for the reaction area were covered with poly (methyl methacrylate) (PMMA), and only the designated area of the MoS_2 single flakes was selectively reacted (Fig. 3b and c). As a result, the expected higher catalytic activity at the edge sites was experimentally demonstrated (Fig. 3c and d).

Therefore, many attempts have been further made to maximize the active site densities of MoS_2 , which has potentially comparable catalytic performances to noble metals. Most intuitively, an interesting approach has been reported that



Fig. 3 (a) On-chip local probe measurement setup for investigating the HER activity of a single MoS_2 flake. (b and c) Optical images of the chips for measuring the HER activities of the (b) $\sim 0.8\ \mu\text{m}^2$ basal plane and (c) $0.3\ \mu\text{m}^2$ basal plane and $0.8\ \mu\text{m}$ edge. (d) Polarization curves of the monolayer MoS_2 basal plane, both with and without edge active sites, in the context of the HER. Reproduced with permission from ref. 14 Copyright 2017 John Wiley and Sons. (e) Scheme of edge-terminated TMDC films with the layers aligned perpendicular to the substrate, maximally exposing the edges of the layers and the corresponding TEM image. Reproduced with permission from ref. 40 Copyright 2013 American Chemical Society. (f) Comparison plot of catalytic activity between flat and nanoscrolls. Reproduced with permission from ref. 75 Copyright 2022 American Chemical Society. (g) Synthesis processes of various morphologies. Reproduced with permission from ref. 76 Copyright 2022 John Wiley and Sons. (h) SEM image of the bulk MoS_2 powder and photos of (i) MoS_2 dispersion with a volume of 60 liters, (j) bare Cu foam and (k) the MoS_2 -based catalyst loaded on Cu foam. Reproduced with permission from ref. 77 Copyright 2020 Springer Nature.

100 mV dec⁻¹, resulting in a significantly improved reaction rate mostly originated from the increased active sites. As revealed, point defects (e.g., vacancies) formed on the basal plane can serve as additional active sites, leading to the pursuit of defect engineering through a mild process rather than complete etching, such as low-energy plasma etching and electrodeposition of metal particles to induce vacancies.^{25,26,28} Fig. 4f illustrates the overall process of maximizing vacancies by sequentially introducing metal electrodeposition following plasma etching.⁴⁸ In the cases where a structure is challenging to etch vertically, vacancies can be formed through a solution process. An unstable and highly oxidizing solution, such as H₂O₂, can be employed for this purpose (Fig. 4g).⁸¹

One of the primary challenges in surface defect engineering *via* etching is the issue of stability, often arising from the unstable flake state. Considering that the presence of chalcogen vacancies in 2D TMDCs can directly impact stability, it is noteworthy that interfaces within 2D materials, such as grain boundaries or homojunctions, exhibit imperfect lattice structures and improved chemical reactivity.^{82–84} Fig. 4h illustrates a process of minimizing vacancies and creating an unstable lattice structure on the basal plane by introducing a third element during crystal synthesis *via* CVD.⁸⁵ The resulting heterojunction flakes exhibit an average hydrogen adsorption free energy of nearly zero (0.26 eV), leading to a significant improvement in the catalyst performance. Furthermore, there are instances where the synthesis of intricate morphology structures, followed by doping, has been employed to maximize the desired effects.^{86–88} For instance, a 3D hierarchical CoP nanoflower, which was synthesized using the metal-organic framework self-sacrifice strategy, can be produced (Fig. 4i).⁸⁶ Additionally, bismuth doping was introduced into the CoP nanoflower, synergistically optimizing the binding free energy of H adatoms on the CoP (202) facet and improving the HER performance. Based on such modification strategies, promising candidates that could genuinely replace noble metal catalysts are increasingly emerging.

2D materials as electrode layers

2D electrodes for loading catalysts: graphene

Graphene, with its high electrical conductivity and high surface area, plays a crucial role as a charge transfer pathway for HER catalysts such as 2D TMDCs and noble metals.^{89–91} Attempts to use high-quality large-area graphene to manufacture a catalytic electrode of a substantial wafer-scale start with finding an effective transfer method. Large-area single-layer graphene grown using CVD is primarily produced through the reaction of methane on copper or nickel (Ni) foil.^{67,92} Subsequently, the substrate needs to be removed and the graphene should be carefully positioned according to the desired location. Polymer-based dry or wet transfer methods have been developed and demonstrated utilizing materials such as PMMA,⁹³ poly(bisphenol A carbonate),^{94,95} polystyrene,⁹⁶ and thermal release tapes.⁹⁷ Among these options, PMMA is the most widely employed due to

its compatibility, light transparency, ease of handling, and adjustable viscosity.^{93,98,99} However, it is known that the polymer is not entirely removed during the transfer process, leaving behind the residue,^{100,101} which is detrimental for charge transfer characteristics that highly suppress the intrinsic characteristics of electrode materials. In addition to the surface contamination, water molecules and electrolyte ions can penetrate between graphene and base substrates, often through edges or small cracks in single-layer graphene during the HER. As the reaction begins and the hydrogen bubbles evolve, the graphene is supposedly separated from the substrate during the reaction.^{102,103}

Producing graphene electrodes for HER that are free from contamination and breakage can be approached in various ways. An improved transfer method, the flipped transfer technique, was developed by incorporating PMMA as an assisting adhesive layer (Fig. 5a).¹⁰³ This method ensures the achievement of a pristine graphene surface free from polymer contaminants. Moreover, the hydrophobic PMMA layer sitting between the graphene electrode and the hydrophilic SiO₂ substrate was optimized to enhance electrochemical reactions by preventing water infiltration and minimizing bubble-induced damage. This configuration is appropriate for electrochemical reactions, as the hydrophobic PMMA layer sitting between the graphene electrode and the hydrophilic SiO₂ substrate prevents water intercalation and eliminates bubble-related damage.

For more practical applications of graphene, efforts have been made to produce large area films on a wafer-scale.¹⁰⁴ Particularly, the solution-based process obtained from LPE has garnered significant interest.^{105,106} LPE is a top-down method suited for mass production and offers cost advantages through a low-temperature process. Using the resulting solution, graphene films can be prepared in various ways, including dip coating,¹⁰⁷ drop casting,¹⁰⁸ spin coating,¹⁰⁹ spray coating,¹¹⁰ and inkjet printing.^{111,112} Fig. 5b compares the effect of a graphene film as a filter for nano-size catalysts with a maximized surface area.⁶⁸ The uniform film was fabricated by simple vacuum filtration without using a binder, thereby preventing problems represented by additional resistance in advance.¹¹³ The cross-section of the film with the catalyst vertically loaded clearly showed the separation of the graphene layer, which effectively acted as a filter (Fig. 5c). Facilitating the movement of electrons from graphene electrodes to catalysts through chemical bonding, rather than simple physical bonding, can be designed. This approach involves the induction of uniform adsorption of metal cations onto the modified graphene sheet through electrostatic attraction, leveraging the strong negative charge on the sheet's surface.¹¹⁴ (Fig. 5d–f) The method enables the production of a practical catalytic electrode utilizing almost 100% of the raw material. This strategy ensures excellent durability and thus provides a new platform for large-scale and high-yield production of graphene nanosheet-based electrocatalysts.

2D electrodes as templates for metallic catalyst growth

The efficiency of catalyst performance depends greatly on how the catalyst is loaded. Typically, catalysts are applied using

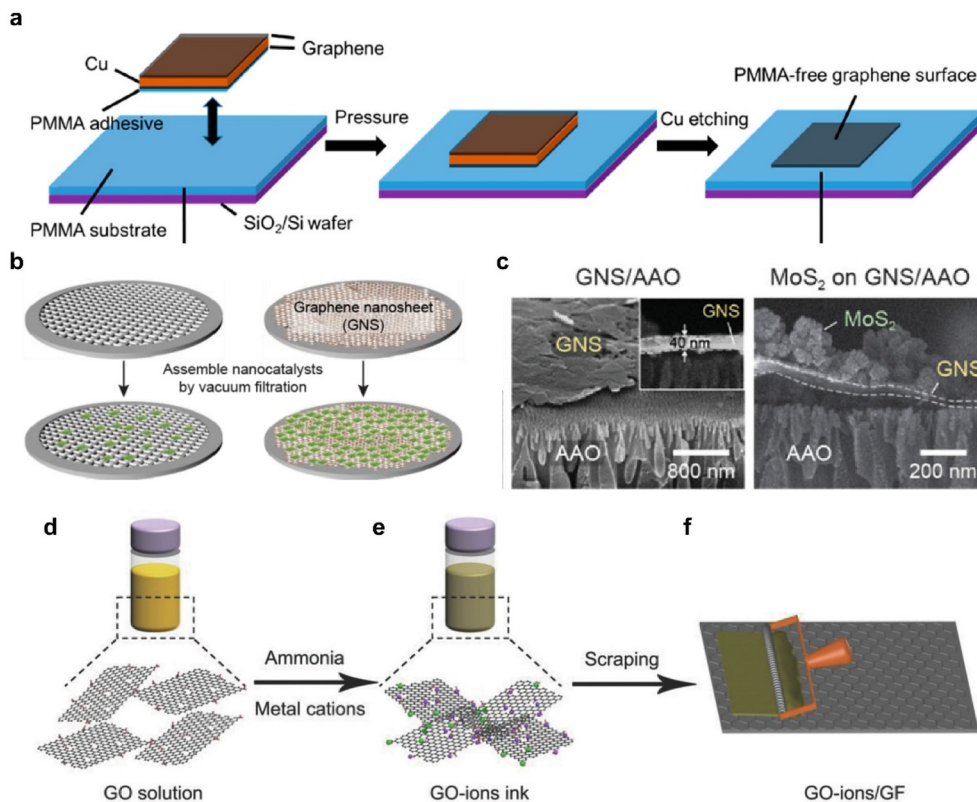


Fig. 5 (a) Schematic illustration of the PMMA-assisted flipped transfer method. In contrast to conventional transfer methods, PMMA is employed as a substrate instead of a sacrificial layer. Reproduced with permission from ref. 103 Copyright 2017 American Chemical Society. (b) Schematic of the graphene nanosheet (GNS)-assisted approach for generating nanocatalyst electrodes. (c) SEM images of the GNS layer and the MoS₂/GNS composite. Reproduced with permission from ref. 68 Copyright 2021 The Authors. (d) Dispersion of a graphene oxide (GO) solution. (e) Addition of ammonia and metal cations to create a GO ion ink. (f) Application of the GO ion ink to the surface of a graphite foil using an automatic film applicator. Reproduced with permission from ref. 114 Copyright 2018 John Wiley and Sons.

ink-based preparation methods onto electrodes, which include drop casting,^{115,116} dip coating,¹¹⁷ and spray casting.¹¹⁶ Some methods involving atomic layer deposition and sputtering have also been suggested.^{118,119} However, ink-based methods often struggle to achieve a uniform thin film, often exhibiting issues like the coffee ring effect and aggregation.^{112,120} Ensuring proper adherence to substrates is another substantial challenge common to the above methods.¹²¹ Given these challenges, research has increasingly focused on cultivating catalysts that form a chemically robust bond with the electrode material, aiming to improve the efficiency. As shown in Fig. 5d–f, 2D vdW materials like graphene offer the advantage of strong and uniform adsorption, even at defect sites along the material's edges and basal plane.¹¹⁴ This characteristic makes these materials excellent templates for catalytic synthesis.

Furthermore, HER catalysts can be directly synthesized on a conductive graphene electrode layer to minimize the contact resistance. In Fig. 6a, Ma *et al.* synthesized MoS₂ on graphene to enhance conductivity for HER application and produced a binder-free, self-standing electrode.¹²² They synthesized MoS₂ through a one-pot solvothermal method, which greatly lowered the reaction rate and succeeded in synthesizing flower-like complex structures. Catalytic metal nanoparticles can also be

synthesized by electrodeposition.^{123,124} Fig. 6b illustrates the synthesized Ni particles near defects in a graphene plane (Fig. 6c).¹²⁴

Furthermore, controlling the morphology¹²³ of catalytic nanoparticles has been considered as an important goal to maximize electrocatalytically active sites and minimize coalescing hydrogen bubbles on the surface. The most critical factor influencing the morphology of particles is the reduction rate of metals.^{125,126} For the successful control over complex morphologies, it is imperative to maintain an extremely low reduction rate, necessitating prolonged reaction times during the overall nucleation and growth stages in this liquid-phase synthesis, which is based on homogeneous nucleation.^{127,128} To mitigate this challenge, the growth mechanism can be adjusted to manipulate the particle morphology by altering the reaction kinetics of Pt species through current modulation.^{129,130} In accordance with the diverse electrical properties of 2D materials, which structurally can serve as an electron carrier, the current flow to the surface can be regulated when employed as a template. In Fig. 6d–g, templates with various current levels were prepared and Pt particles were electrochemically deposited.³³ Intriguingly, the shape and area density of the Pt nanoparticles exhibited significant variations depending on the



Fig. 6 (a) Scheme of the MoS₂NF/rGO paper fabrication process and SEM images illustrating the cross-section and top view of MoS₂NF/rGO. Reproduced with permission from ref. 122 Copyright 2014 Royal Society of Chemistry. (b) Preparation process of the Ni-rGO catalyst. (c) TEM image of a single Ni particle supported on a rGO sheet and HRTEM images of the rGO sheet. The inset shows a closer view. Reproduced with permission from ref. 124 Copyright 2017 Elsevier. (d–g) Top-view SEM images of Pt nanoparticles grown on CP, graphene, eMoS₂, and sMoS₂ at 0.5 V. Reproduced with permission from ref. 33 Copyright 2022 The Authors. (h) Schematic representation of ultra-small Pt particles generated *in situ* on the surface of a PtSe₂ nanosheet with Se vacancies. (i) The structural model used in density functional theory (DFT) calculations: a PtSe₂ nanosheet decorated with a Pt₁₃ cluster. Colored circles mark the active sites and V_n indicates the inclusion of *n* Se vacancies. (j) TEM images of PtSe₂/Pt. (k) Magnified TEM image of PtSe₂/Pt. Reproduced with permission from ref. 136 Copyright 2021 John Wiley and Sons.

growth template used. These template-dependent differences in the Pt nanoparticle morphology can be attributed to variations in the reduction rate of Pt ions for each template.^{129–131} Templates with higher resistance result in a slower reduction of Pt ions compared to the diffusion transport of the precursor, preventing the formation of a blank layer around the nucleus and maintaining an appropriate balance.^{129–131} Under such conditions, protrusions, such as edges or corners, are exposed to the precursor more than the particle core, promoting faster growth and leading to the formation of a flower-like structure.^{129–131} This groundbreaking strategy enables precise shape control through a brief electrochemical reaction and holds immense potential for various applications, as it can effectively maximize the catalyst's active surface area.

Lastly, 2D noble metal dichalcogenides (NMDs), which have attracted attention for their high carrier mobility and stability, can be another candidate for the strong interaction between particles and templates.^{132,133} For transition metal com-

pounds, d-electrons define the phase structure and coordination environment, and the resulting interaction strength with active species.¹³⁴ In particular, the d-band, which represents electron states associated with the d-orbitals of transition metals, exhibits better stability because it is highly occupied in NMDs, leading to strong metal adsorption on the surface.¹³⁵ Pt, a representative catalyst for HER, can construct a heterointerface with PtSe₂ through an electrochemical reduction method (Fig. 6h).¹³⁶ At this time, it was found that a continuous Se vacancy was induced around the PtSe₂/Pt boundary to strengthen the connection with particles (Fig. 6i). Large amounts of Pt particles at the edges of PtSe₂ were observed, demonstrating the successful synthesis, and leading to good catalytic performance (Fig. 6j and k).

2D materials as interlayers

In recent studies, researchers have focused on exploring the interface layer between the catalyst and the electrode, which

plays a crucial role in facilitating charge transfer between the catalyst and the electrode, while also imparting additional functionalities. One of the key factors under investigation is the Schottky barrier, which represents the tunnel barrier encountered by electrons at the interface between the semiconductor catalyst and the conductive support.¹³⁷ This barrier can pose challenges to catalytic activity, primarily due to inefficient charge injection. From this standpoint, research has been conducted to accurately understand how interlayers influence the control of HER activity, considering the fact that adjacent support materials can alter various factors affecting the HER, such as ΔG_H and the height of the Schottky barrier.^{137,138} Fig. 7a shows a photograph of a microcell constructed using semi-metallic WTe₂ as an interlayer between MoS₂ and Au electrodes.¹³⁹ Micromechanically-exfoliated flakes were stacked, and HER activities were elucidated to evaluate the influence of the interlayer without being affected by other factors. As a result, the relatively close Fermi level alignment between MoS₂ and WTe₂ resulted in a lower Schottky barrier, which could facilitate efficient electron injection to improve HER activity (Fig. 7b).^{137,138} Fig. 7c shows that a perfectly stacked structure exhibits higher catalytic activity in common than a structure

simply connected to MoS₂ and WTe₂, which emphasizes the need for an interlayer structure.¹³⁹

In addition, Fig. 7d illustrates how electrons reach the HER reaction within the context of energy band alignment when an interlayer is introduced, also providing a potential applicability as a photocatalyst.^{50,140} The band energies of various 2D materials for water redox potentials have been extensively investigated, and 2D photocatalysts can be selected for water reduction or oxidation reactions based on valence band maximum (VBM) and conduction band minimum (CBM) locations.¹⁴¹ As depicted in Fig. 7e, the creation of a p-n junction structure is an essential process for implementing the photocatalyst.²⁸ In the energy level diagram, p-type MnCo-CH possesses only the HER potential, while NiFe-OH features an overlapping energy level accommodating both electrons and the OER potential (Fig. 7f).²⁸ When these two materials come into contact by forming a p-n junction, an adjustment in energy levels occurs, resulting in distinct catalytic behaviours for OER and HER within the hybrid structure (Fig. 7f). Electrons flow from NiFe-OH to MnCo-CH and recombine with holes until reaching Fermi level equilibrium, causing band bending.²⁸ This bending increases the gap between NiFe-OH's valence band and the OER potential, facilitating

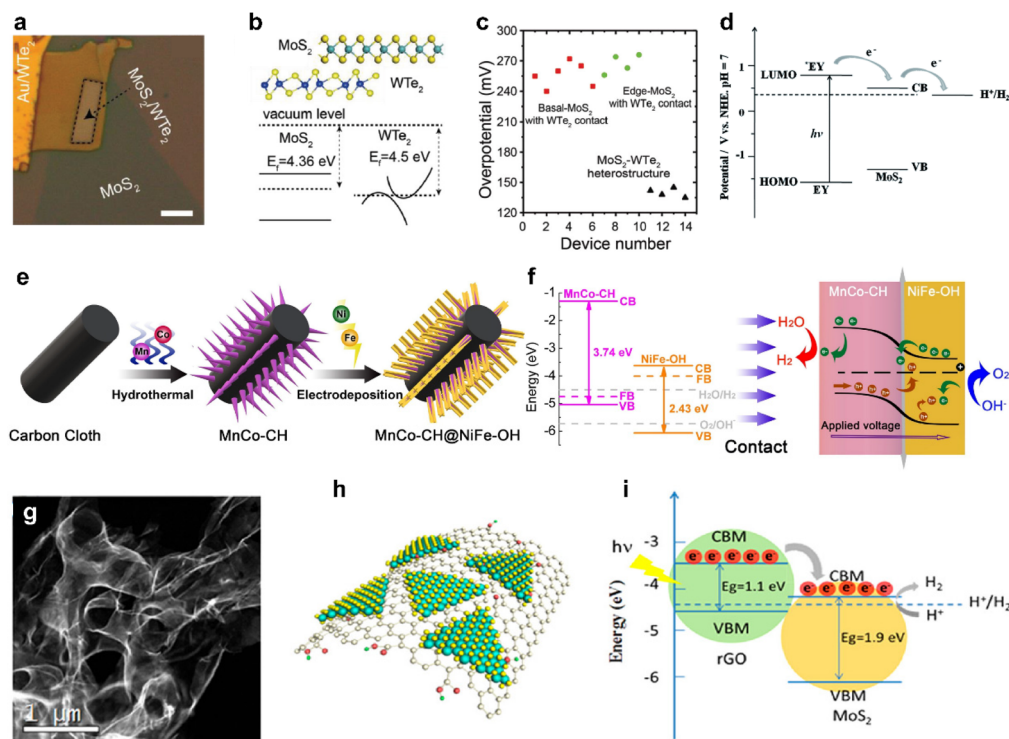


Fig. 7 (a) Optical image of the MoS₂-WTe₂ heterostructure. (b) Scheme of the MoS₂-WTe₂ heterostructure and the band alignment at the MoS₂-WTe₂ interface. (c) Comparison plot of overpotentials. Reproduced with permission from ref. 139 Copyright 2019 John Wiley and Sons. (d) Proposed mechanism for photocatalytic H₂ evolution under visible light irradiation and the energy level for the EY-MoS₂-TEOA system. Reproduced with permission from ref. 140 Copyright 2016 Royal Society of Chemistry. (e) Synthetic process of the MnCo-CH@NiFe-OH p-n junction. (f) Energy diagrams illustrating MnCo-CH and NiFe-OH (left) and the MnCo-CH@NiFe-OH p-n junction (right). Reproduced with permission from ref. 28 Copyright 2021 Elsevier. (g) STEM image of 3D nanoporous MoS₂@rGO. (h) Schematic representation of MoS₂@rGO. (i) Schematic illustration of band alignment and the mechanism of photoexcited charge transfer in MoS₂@rGO. Reproduced with permission from ref. 144 Copyright 2018 American Chemical Society.

Table 1 HER performance of 2D materials-based catalytic electrodes

Sample	Synthesis method	Electrolyte	Overpotential	Tafel slope	Remark	Ref.
2H-MoS ₂ basal plane	Chemical vapor deposition	0.5 M H ₂ SO ₄	425 mV(@10 mA cm ⁻²)	109 mV dec ⁻¹	—	14
2H-MoS ₂ Mo edge	Chemical vapor deposition	0.5 M H ₂ SO ₄	201 mV(@10 mA cm ⁻²)	68 mV dec ⁻¹	Active site	14
2H-MoS ₂ on graphene	Liquid phase exfoliation	1 M KOH	486 mV(@10 mA cm ⁻²)	97 mV dec ⁻¹	Wafer-scale/graphene electrode	68
1T'-MoS ₂ basal plane	Chemical vapor deposition	0.5 M H ₂ SO ₄	356 mV(@10 mA cm ⁻²)	84 mV dec ⁻¹	—	14
1T'-MoS ₂ Mo edge	Chemical vapor deposition	0.5 M H ₂ SO ₄	77 mV(@10 mA cm ⁻²)	60 mV dec ⁻¹	Active site	80
1H-MoS ₂ edge site	Chemical vapor deposition	0.5 M H ₂ SO ₄	940 mV(@30 mA cm ⁻²)	130 mV dec ⁻¹	Active site	137
1H-MoS ₂ terrace	Chemical vapor deposition	0.5 M H ₂ SO ₄	1060 mV(@30 mA cm ⁻²)	130 mV dec ⁻¹	—	136
WTe ₂ basal plane	Mechanical exfoliation	0.5 M H ₂ SO ₄	390 mV(@10 mA cm ⁻²)	141 mV dec ⁻¹	Active site	23
WTe ₂ (010) edge	Mechanical exfoliation	0.5 M H ₂ SO ₄	350 mV(@10 mA cm ⁻²)	135 mV dec ⁻¹	Active site	24
WTe ₂ (100) edge	Mechanical exfoliation	0.5 M H ₂ SO ₄	320 mV(@10 mA cm ⁻²)	110 mV dec ⁻¹	Active site	33
PtSe ₂	Solid-state synthesis	1 M KOH	294 mV(@10 mA cm ⁻²)	147 mV dec ⁻¹	—	124
Pt on N-doped graphene	Sacrificial template method	0.5 M H ₂ SO ₄	30 mV(@10 mA cm ⁻²)	28 mV dec ⁻¹	Graphene electrode	81
Single atom Pt on graphene	Microwave-assisted synthesis	0.5 M H ₂ SO ₄	12 mV(@10 mA cm ⁻²)	29 mV dec ⁻¹	Graphene electrode	48
Pt-deposited MoS ₂	Electrodeposition of Pt	1 M KOH	74 mV(@10 mA cm ⁻²)	53 mV dec ⁻¹	Surface modification	136
Ni-rGO _{1.0} /Ni foam	Supergavity electrodeposition	1 M NaOH	36 mV(@10 mA cm ⁻²)	77 mV dec ⁻¹	Surface modification	139
Etched MoS ₂ basal plane	Steam vapor etched @700 °C	0.5 M H ₂ SO ₄	350 mV(@10 mA cm ⁻²)	96 mV dec ⁻¹	Surface modification	85
Etched MoS ₂ nanosheets	Immersed in 5 mol L ⁻¹ H ₂ O ₂ solution	0.5 M H ₂ SO ₄	131 mV(@10 mA cm ⁻²)	48 mV dec ⁻¹	Surface modification	136
RIE-treated 1T'-PtSe ₂	Ar plasma treatment	1 M KOH	198 mV(@10 mA cm ⁻²)	53 mV dec ⁻¹	Surface modification	139
Pt-deposited 1T'-PtSe ₂	Electrodeposition of Pt	1 M KOH	27 mV(@10 mA cm ⁻²)	40 mV dec ⁻¹	Heterostructure	85
Pt/PtSe ₂	Electrochemical reduction method	1 M KOH	42 mV(@10 mA cm ⁻²)	58 mV dec ⁻¹	Heterostructure	85
2H-MoS ₂ /WTe ₂	—	0.5 M H ₂ SO ₄	150 mV(@10 mA cm ⁻²)	—	—	—
1T'-Re _x Mo _{1-x} S ₂ /2H-MoS ₂	—	0.5 M H ₂ SO ₄	84 mV(@10 mA cm ⁻²)	—	—	—

electron flow and a positive charge on NiFe-OH, enhancing OH⁻ movement and promoting the OER.²⁸ Simultaneously, it widens the gap between MnCo-CH's conduction band and the HER potential, boosting electron mobility for H⁺ reduction to H₂.²⁸

Photocatalytic design can be realized using monolayer 2D materials.^{142,143} For instance, a heterojunction between p-type reduced graphene oxide (rGO) and n-type MoS₂, synthesized *via* CVD, provides a suitable example (Fig. 7g).¹⁴⁴ In this configuration, the CBM position of rGO is lower than that of MoS₂, with the energy level of the HER situated below the CBM of MoS₂. Consequently, this band alignment creates type-II (staggered) heterojunction structures that facilitate effective charge transfer upon photoexcitation. When exposed to visible light, electrons are excited from the VBM of rGO to the CBM. The designed CBM offset between MoS₂ and rGO promotes efficient photoexcited electron transfer and facilitates the separation of holes and electrons, resulting in an electron-rich MoS₂ layer conducive to the HER.¹⁴⁴ Additionally, the photo-generated electrons contribute to a positive bias during the HER process, further enhancing the catalytic performance of the MoS₂@rGO catalyst.¹⁴⁴

Conclusions and outlook

Scalable synthesis of electrocatalytic-grade 2D materials with desired electronic properties to implement into catalytic electrodes by understanding the roles of each part including a catalytic layer, an electrode, and an interfacial layer can be an important first step towards enhancing the 2D material-based HER performance. Although considerable theoretical and empirical studies of 2D materials have been reported for their use in a catalytic layer and an electrode, investigations of inter-layer applications remain relatively challenging due to the lack of appropriate processing methods for scalable p-n heterostructure formation. Heterostructures have shown the potential to enhance both the HER performance and functionality, but their accessibility has been limited by often involving sequential synthesis processes.^{28,145} However, the structural advantages of vertical stacking in 2D materials can be a key in addressing this issue. Furthermore, their wide range of electrical properties offers extensive choices for structure design based on energy bands. The excellence of 2D materials used as catalysts and electrode layers is summarized in Table 1. The HER performance of active sites in pristine 2D materials indicates the necessity of catalyst surface modification, and the modified catalysts exhibit outstanding characteristics. Furthermore, the adoption of graphene as an electrode layer is proven to be feasible by using Pt composites. However, due to the limited research on the role of functional 2D materials in enhancing the performance of catalytic electrodes as inter-layers, only a few cases could be added. Improving synthesis and large-area film fabrication methods for effective vertical stacking can broaden the scope of functionality, ultimately leading to the development of more advanced 2D catalysts.

This departure from the conventional paradigm focused on catalytic surfaces provides new insights into increasing catalytic activity in future research.

Conflicts of interest

There are no conflicts to declare.

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