

Cite this: *RSC Sustainability*, 2023, 1, 11

From rocks to bioactive compounds: a journey through the global P(v) organophosphorus industry and its sustainability

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This tutorial review aims at presenting the usefulness and the diversity of organophosphorus compounds within the context of a more sustainable phosphorus chemistry. Three topics will be emphasized while reflecting on green chemistry: firstly, a short sum-up of the phosphorus resources and their industrial and environmental ties will be given to contextualize the framework of this review. The rich nomenclature, properties, and applications of organophosphorus compounds will then be highlighted, emphasizing P(v) comparatively to P(III) ones. Finally, a brief review of the current synthetic routes towards organophosphorus molecules and their inherent challenges will be described before presenting a few promising alternatives.

Received 13th August 2022
Accepted 14th October 2022

DOI: 10.1039/d2su00015f

rsc.li/rscsus

Sustainability spotlight

With the growing need for the establishment of a greener chemistry, a closer look at the current organophosphorus chemistry field is required. Most synthetic phosphorus compounds are indirectly obtained through the functionalization of extracted phosphate rocks, a finite resource which is asymmetrically spread around the globe. The overreliance on energy-costly processes, on transformations generating toxic by-products, and the use of the same hazardous intermediates within the organophosphorus industry will be described. In line with the sustainable development goals 12 and 15 of the United Nations about sustainable production patterns & industrialization, the proposed tutorial review showcases an accessible and up-to-date description of the various applications of P(v) molecules and their current industrial synthetic routes, before investigating the inherent sustainability challenges of the field and highlighting more sustainable prospective alternatives.

1. An overview of phosphorus processes

1.1 The universality of phosphorus in life-essential structures

While organic chemistry focuses mainly on the chemical behavior of carbon, hydrogen, and a few period 2 elements such as nitrogen and oxygen, and elements from the halogen group, the chemistry of sulfur and phosphorus remains somewhat marginal despite their abundance in natural patterns. Indeed, phosphorus, the predominant element examined in the present review, is omnipresent in biomolecules and essential to life. Two remarkable examples can be found in biochemistry, thus highlighting its significance: one example sees phosphorus playing a role in the architecture of information encoding biopolymers. The second shows phosphorus as the universal mediator for energy transfer in life processes. Contrary to sulfur which exists directly in the peptide structure, phosphorus is present as phosphates in the skeleton of polynucleotides such

as DNA or RNA. Each nucleoside is linked to another unit through a 3'-5' phosphodiester linkage on the sugar rings (Fig. 1a). Being anionic at pH = 7, phosphates confer the overall biopolymer architecture a good water-solubility and thus physiological compatibility, which would otherwise be lacking due to the relative hydrophobicity of the nitrogenous base.

Furthermore, the negative charges allow for specific proteins to bind to the DNA strands easily. After elucidation of the DNA structure by Watson, Crick, Wilkins, and Franklin,¹ it was shown that this sugar-phosphate backbone twists around itself



Fig. 1 Common phosphate containing biomolecules. (a) Guanine nucleotide unit (b) adenosine triphosphate.

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Scheme 1 Oxidation of white phosphorus.

to form the well-known double helices structure. Phosphorus is also present in adenosine triphosphate (ATP), the molecular unit of energy present in all known forms of life (Fig. 1b). Synthesized during cellular respiration, photosynthesis, or fermentation, ATP is used to fuel multiple metabolic reactions, transport macromolecules across the cellular membrane, DNA and RNA synthesis, muscle contractions, *etc.* The chemical energy transferred during those processes relies on the hydrolysis of the terminal phosphate yielding one adenosine diphosphate unit, one free phosphate anion, and $7.3 \text{ kcal mol}^{-1}$ per ATP molecule consumed.² Phosphorylation of proteins – understood in biochemistry by adding a phosphate unit onto an alcohol functional group – is also the most abundant modification and has a prominent role in activating, deactivating, or modifying the role of enzymes.³ Phosphorus can also be found in cell membranes as phospholipids or inorganic apatite in bones and teeth enamel.

This entire section will give a brief timeline of phosphorus discoveries. A short overlook of the global phosphorus resources available and the industrial processes will then be given. Finally, we will briefly investigate the ties linking phosphorus and the environment.

1.2 A brief history of phosphorus discoveries

Despite its universal role in life processes, elemental phosphorus was isolated in 1669 by German alchemist Hennig Brand during his efforts to find the philosopher's stone.⁴ Brand



Fig. 2 Timeline of the discovery of significant phosphorus-containing compounds.

obtained a liquid that instantly gave off fumes and burst into flames when distilling urine. Moreover, this substance could solidify and glow in the dark, which is why he coined it phosphorus, from 'Φωσφόρος' which means light-bearer in Greek. P₄, in the white phosphorus form, was thus well-known from the beginning as a pyrophoric substance, as it readily converts to phosphorus pentoxide P₂O₅ in the presence of oxygen from the air (Scheme 1).

After its discovery, phosphorus was soon found to be present in plants, bones, and minerals such as pyromorphite, and by the late 18th century, PH₃ was prepared by Gengembre,⁵ followed by Gay-Lussac and Thénard, who prepared PCl₃ and PCl₅ in 1808 (Fig. 2).⁶ By the 1830s, white phosphorus was used in France and England in striking matches, unfortunately leading



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to many fatalities amongst factory workers.^{7,8} In 1842, Lawes and Murray patented in Great Britain the first technique for synthesizing P-fertilizer coined 'superphosphate' using sulfuric acid and bones.⁹ However, even though bones and later guano had replaced urine as the primary source of phosphorus, the discovery of phosphorus-rich mineral deposits by the late 19th century changed the face of the blooming phosphorus economy. Those phosphate rocks deposits, such as in Suffolk in England, were readily exploited to satisfy the increasing demand for fertilizers. Later discovery of the wet process to obtain phosphoric acid and the invention of the electric furnace in 1888 by Readman paved the way for more efficient industrialization and commercialization of phosphorus compounds.¹⁰ As for metal-phosphorus complexes, the pioneering works of Rose, Hofmann, Cahours, and Gal during the second half of the 19th century foreshadowed the extensive use of transition-metal phosphine complexes in modern catalysis.^{11–14} Although Lassaigne had already prepared alkyl phosphates in 1820,¹⁵ it was both Arbuzov and Michaelis,^{16,17} now primarily considered as the founders of modern organophosphorus chemistry, who became well-known for developing coupling methodologies to construct C–P bond, thus creating new tools for the introduction of phosphorus moieties in organic synthesis. The diverse use of inorganic phosphates came about gradually from the 1930s until the 1950s, when they became frequently employed for toothpaste, surfactants, water softening and food technology.¹⁸ During this era, a German chemist called Gerhard Schrader also patented the first organophosphorus insecticides: hexaethyltetraphosphate (HETP) and tetraethylphosphate (TEPP).¹⁹ However, his discoveries are also closely associated with the most historically pernicious use of phosphorus: chemical weaponry. The use of elemental phosphorus as incendiary bullets dates back from World War I. Nevertheless, it was Schrader who developed diverse nerve agents, amongst which the infamous Tabun and Sarin gas,²⁰ which are now classified as weapons of mass destruction and formally banned under the Chemical Weapons Convention of 1993.²¹ Although elemental phosphorus has now been restricted in the case of incendiary devices, white phosphorus munitions are still in usage to produce smoke and camouflage movement. The effects of white phosphorus munitions remain still heavily dangerous, causing significant burn injuries and respiratory damages.

1.3 Phosphate rocks: a scarce resource

Phosphorus is not a rare element on the surface of the earth. It is thought to make up between 1000 to 1300 ppm by mass of the Earth's continental crust, and as such, is more abundant than chlorine or sulfur, or even carbon.^{22–25} However, most of the exploitable phosphorus sources reside in sedimentary phosphate rocks, typically fluorapatite $\text{Ca}_5(\text{PO}_4)_3\text{F}$. The availability of those phosphate rocks is finite, and various estimation of the longevity of the world's reserve varies between complete depletion coming about within the next 70 to 400 years.²⁶ The continuously changing assessment over this controversial issue can be explained by relying on limited but ever-changing new available information and considering the future economic and

Table 1 Reserves and production of phosphate rocks in 2019 by country

Country	Reserves of P-rocks in 2019 (ton × 10 ⁶)	Country	Production in 2019 (ton × 10 ⁶)
Morocco	50 000	China	110
China	3200	Morocco	36
Algeria	2200	USA	23
Syria	1800	Russia	14
Brazil	1700	Jordan	8
Saudi Arabia	1400	Saudi Arabia	6.2
South Africa	1400	Vietnam	5.5
Egypt	1300	Brazil	5.3
Australia	1200	Egypt	5
USA	1000	Peru	3.7
Finland	1000	Israel	3.5
Jordan	1000	Tunisia	3

technological dynamics of exploitation and evaluation.²⁷ The exact magnitude of the scarcity of phosphorus is still far from settled. However, a concept known as peak phosphorus, the point in time when the maximum production rate of phosphorus will be reached, has emerged as an analogy to the concept of peak oil.²⁸ Most of the commercial reserves of phosphate rocks (94%) are split between only 11 countries (Table 1), and 70% of the reserves lies solely within Morocco/Western Sahara. In 2015, the global extraction capacity of phosphate rocks reached 223×10^6 tons, with China, the US, and Morocco representing 70% of the extraction alone.²⁹

1.4 The phosphorus industry and processes

It has been estimated that over 90% of the phosphate rock exploited is converted to phosphoric acid according to the 'wet process' (Scheme 2). The method consists of the digestion of fluorapatite minerals by concentrated sulfuric acid. Calcium sulfate which is yielded as a by-product (about 4–6 tons produced per ton of phosphoric acid) is later precipitated and filtered out, and the wet phosphoric acid obtained is purified and concentrated.³⁰

It is estimated that this process accounts for more than half of the global supply of sulfur taking the form of sulfuric acid, thus showing how tightly the phosphorus and sulfur industries are interdependent.³¹ Most of the phosphoric acid produced is used for the fertilizer industry. The 4 commonly sold phosphorus fertilizers and their production in 2016 are shown above (Fig. 3).³² The remaining part of phosphate rocks can be converted to molecular phosphorus P_4 for subsequent reactions. Historically, phosphoric acid was also synthesized through this process by oxidizing the white phosphorus obtained to afford phosphorus pentoxide, the anhydride version of phosphoric acid. However, because the reaction is strongly endothermic,



Scheme 2 Conversion of phosphate rock to orthophosphoric acid through the wet process.



		formula	2016 production (ton x 10 ⁶)
MAP	monoammonium phosphate	NH ₄ H ₂ PO ₄	29
DAP	diammonium phosphate	(NH ₄) ₂ HPO ₄	34
TSP	triple superphosphate	Ca(H ₂ PO ₄) ₂ ·H ₂ O from P ₄	6
SSP	single superphosphate	Ca(H ₂ PO ₄) ₂ ·H ₂ O from apatite	25

Fig. 3 Commonly sold phosphorus-containing fertilizers.

this thermal process requires electric furnaces, which are highly energy intensive. These issues caused the development of the wet process, which circumvented the need to go through a white



Scheme 3 Conversion of phosphate rocks to white phosphorus through the Wöhler process.

phosphorus intermediate to obtain phosphoric acid. Nevertheless, this thermal process (coined the Wöhler process) remains the most efficient pathway to produce molecular phosphorus. An electric furnace heating at around 1450 °C is fed the three raw materials: phosphate rocks that are either sintered or pelletized, dried metallurgical coke for the carbon source, and silica deposits such as pebbles, gravel, or coarse sand (Scheme 3).³³

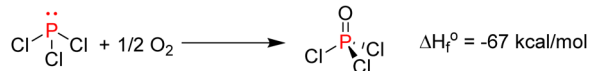
The P-gas produced are then condensed and washed with water. The white phosphorus afforded is then stored under-water to prevent it from self-ignition. Even though the wet process mostly supplanted the thermal pathway, around 83% of the P₄ obtained is still converted to phosphoric acid through combustion.³³ This phosphoric acid has the advantage of being purer than the one obtained through the wet process and can be used accordingly for more sensitive applications. Only about 12% of the white P₄ produced is converted into fine chemicals and heavy phosphorus compounds, including PCl₃, PCl₅, and P₄S₁₀ with the remaining being used for all other industrial phosphorus species (Fig. 4).³⁴

One of the most important classes of inorganic compounds synthesized from elemental phosphorus is phosphorus halides.



Fig. 4 Overview of the phosphorus industry and application.



a. Synthesis of phosphorus trichloride**b. Synthesis of phosphorus pentachloride****c. Synthesis of phosphoryl chloride**

Scheme 4 Synthesis of organophosphorus chloride species.

Phosphorus trichloride is produced by the combustion of either molten white phosphorus or red phosphorus with gaseous chlorine (Scheme 4a). The high heat afforded by this exothermic reaction keeps the process at a boiling point and allows for phosphorus trichloride to be collected by distillation.³⁵ The world production surpasses 300 000 tons per year,³⁶ and it is the most critical building block for the synthesis of many valuable organophosphorus species, as the final section of this review will explain in further detail. Phosphorus pentachloride is another halogenated phosphorus compound of relative importance, and it is produced by reacting phosphorus trichloride with an excess of chlorine gas (Scheme 4b). As a point of comparison, around 10 000 tons of PCl_5 were produced in 2000 worldwide.³⁷ It is mainly used to synthesize acyl chloride or the chlorination of alcohol as PCl_5 is an efficient chlorinating agent thanks to the enhanced electrophilicity of its pentavalent phosphorus center. Finally, phosphoryl chloride, another precursor used to produce phosphate esters, can be synthesized by oxidizing PCl_3 with air or pure oxygen (Scheme 4c). Overall, those halogenated phosphorus species are particularly crucial in the bigger picture as they are precursors for the synthesis of organophosphorus molecules.³⁸ Nevertheless, their production and handling are particularly problematic due to their high toxicity and hazardous characteristics.

1.5. The phosphorus cycle, human-caused perturbations, and environmental concerns

1.5.1. The natural phosphorus cycle. Before delving into the impressive variety of organophosphorus species and their uses, a brief overlook of the interlinks between phosphorus and the Earth's environment is needed to understand the sustainable processes at stake. As mentioned earlier and despite the controversy over the actual longevity of the reserves, phosphorus resources we rely on are non-renewable, and some of the industrial processes used to synthesize organophosphorus need to be revised in terms of sustainability. In addition to those ongoing anthropogenic issues, it is essential to envisage the outlook of phosphorus chemistry holistically. In analogy with carbon or nitrogen, phosphorus-based chemicals undergo

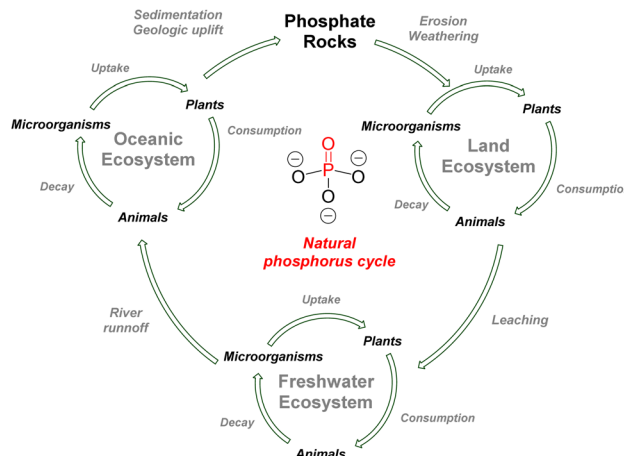


Fig. 5 Simplified schematic of the natural phosphorus cycle on Earth.

a biogeochemical life cycle on Earth (Fig. 5). The singularity of the phosphorus cycle is the absence of a significant gas component, as PH_3 natural production is only localized and inconsequential. As previously stated, most of the phosphorus available on the surface originates from apatite rocks. Due to weathering and natural erosion, those minerals will slowly release phosphate anions into the environment, into the soil, and rivers. After the runoff, phosphate anions (mainly orthophosphate PO_4^{3-}) will then follow fluvial waters, be delivered into the ocean, and will finally precipitate as sedimentary rocks, which will subsequently be buried in the depth of the seabed. The tectonic plaque uplift will require millions of years to make those sedimentary minerals available again on land to close the loop of this slow biogeochemical cycle. On top of this primary mineral cycle following the geological timescale are additional and much faster organic phosphorus cycles depending on ecosystems and life processes. On land, this system is a loop between the dissolved phosphate ions present in the soil, plants, and microorganisms using those ions for their growth and the synthesis of organic compounds, animals consuming those plants and incorporating phosphorus in their tissues, and finally, their excretion or after-death decay returning phosphorus to the soil. Similar systems exist in freshwater or oceanic biotas with phosphorus transfer from solubilized ions to diverse living organisms and the subsequent return of phosphorus to the water according to their decay or waste.

1.5.2. Eutrophication: the danger of leaching excess phosphate. The global equilibrium binding these interdependent cycles remains precarious. A simple increase in the delivery of phosphate ions into the ocean through riverine runoff can have devastating consequences: the sudden surge of phosphorus in the ocean is known to trigger mass extinctions. A high concentration of phosphate ions favors the excessive proliferation of algae in bodies of water. Although temporarily increasing the biomass, this process called eutrophication has been linked to the death of ecosystems by several factors: blocking sunlight below the algal layer, increasing dissolved CO_2 concentration and water acidity, depletion of dissolved



oxygen, and loss of biodiversity. For example, eutrophication caused by perturbations in the oceanic phosphorus cycle has been recorded during the Late Devonian extinction, which caused the elimination of 75% of all species.³⁹

1.5.3. Human impact on the phosphorus cycle: issues and prospective solutions. The massive anthropogenic release of otherwise unavailable phosphorus into the environment caused by the extensive use of fertilizers and detergents, the production of manure, and the discharge of human-generated waste is problematic. It has already caused local eutrophication events in North American great lakes or the Baltic sea.^{40–42} By now, a total of $1.1 \times 10^2 T_g$ of phosphorus has been extracted since the 1840s and injected suddenly in the cycle. The phosphorus flux from the riverine runoffs has already increased over two and a half times from its background state (4.5 to $12.6 T_g \text{ year}^{-1}$).²⁶ As such, Rockström proposed that disturbances of the P-cycle must be listed as one of the “9 planetary boundaries” not to transgress to keep human activities from causing unacceptable change.⁴³ The risks are twofold: ecological and environmental in the short term and food security in the long term. Phosphorus resources need to be managed carefully: efficiency improvement for fertilizers and reducing their overuse is to be implemented. Recycling technologies and the recovering of phosphorus through wastewater, mining waste, manure, sludge, and food residues are already developing.^{26,27,31} An excellent review recently published by Slootweg maps out the different alternatives for phosphorus recovery and recycling within the context of a more circular economy.⁴⁴ The authors conclude that although present economic barriers still hinder large-scale phosphorus recovery (for instance, in the form of struvite recovery processes) because of high implementation costs compared to a relatively cheap price of phosphate rocks, a future surge in their prices combined with

a stronger social drive and policymakers incentives could push the industry to finally “close the loop” of the phosphorus cycle.

2. Organophosphorus compounds: chemical properties and applications

This section will be detailed on organophosphorus compounds: their nomenclature, chemical properties, and applications in various fields. In addition, an emphasis will be given on sustainability issues linked to the industry, and a focus on pentavalent organophosphorus compounds will be made.

2.1. Nomenclature and classification

2.1.1. An intimidating nomenclature. The broadest definition of organophosphorus molecules includes all organic compounds containing phosphorus and is not necessarily reduced to structures only containing P–C bonds, which can specifically be called carbophosphorus. Before exploring their wide range of applications and their synthesis, a look at the complex nomenclature of organophosphorus is needed. The profuse number of organophosphorus species and the multiple names available for the same structure sometimes led to the misnaming of molecules or transformations described in a certain number of publications. The confusion is common, and this section is here to establish the nomenclature used in the present review and act as a tutorial for the reader. Although the overall nomenclature can seem quite intimidating, a straightforward way to directly categorize the most common organophosphorus species is to split them according to trivalent and pentavalent species (Fig. 6). The P(III) compounds bear 3 single bonds and 2 π electrons as a lone pair in an overall trigonal pyramid geometry. This lone pair gives P(III) species

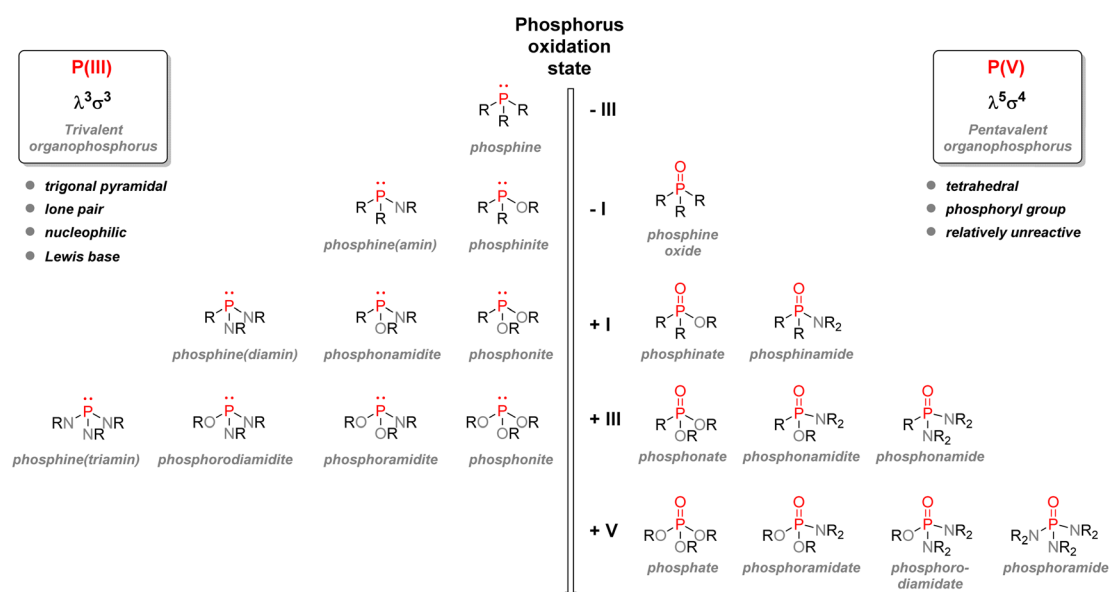


Fig. 6 Nomenclature of organophosphorus compounds: trivalent and pentavalent species sorted by oxidation number.



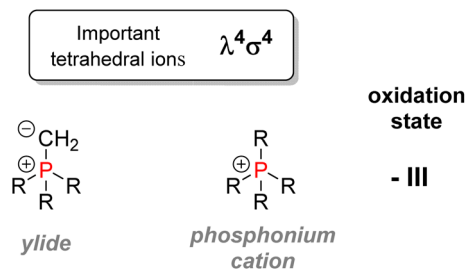


Fig. 7 Important tetrahedral organophosphorus ions.

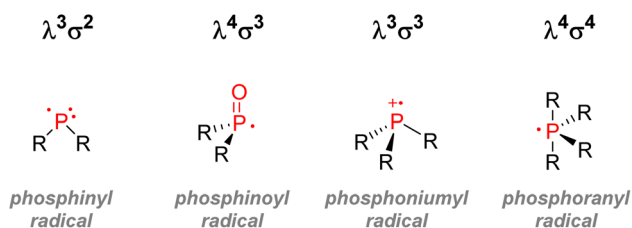


Fig. 8 Important organophosphorus radicals.

nucleophilic properties and Lewis base characteristics. They can also be used as reductants. Pentavalent organophosphorus are usually poorly reactive and comprise 3 single bonds and a phosphoryl bond P=O or phosphorothioate P=S bond arranged within a tetrahedral geometry. Another way to classify

organophosphorus compounds is according to the numbers of P–C, P–O, or P–N bonds or according to the phosphorus oxidation state: it goes from –III in the case of phosphines PR₃ or +V in the case of phosphates P(O)(OR)₃.

2.1.2. Common ionic organophosphorus species. Aside from those commonly encountered molecules, a few essential tetrahedral ions should be added: ylides are zwitterionic species bearing a positive charge on the phosphorus and a carbanion, and phosphoniums are tetrahedral organophosphorus with a positive charge on the P-center and often serve a reactive intermediate (Fig. 7).

2.1.3. Common radical organophosphorus species. Besides ionic phosphorus, P-centered radicals are also crucial reactive intermediates (Fig. 8). The first ones are phosphinyl radicals which have a disubstituted bent-like geometry with a π -character SOMO. Phosphinoyl radicals (sometimes called phosphonyl radicals) are σ -radicals bearing a trigonal pyramidal geometry with two substituents and a phosphoryl bond; their reactivity has already been extensively studied. Phosphoniumyl radicals originate from P(III) species which got oxidized through SET to afford a trigonal pyramidal σ -radical cation. Phosphoranyl radicals are more diverse in their geometries, adopting either a pseudo-trigonal bipyramidal geometry or a tetragonal structure.⁴⁵

2.1.4. Valence nomenclature and less common organophosphorus compounds. Most organophosphorus molecules encountered in this review will belong to the P(III)/P(V) categories exposed in the scheme above, but another classical yet

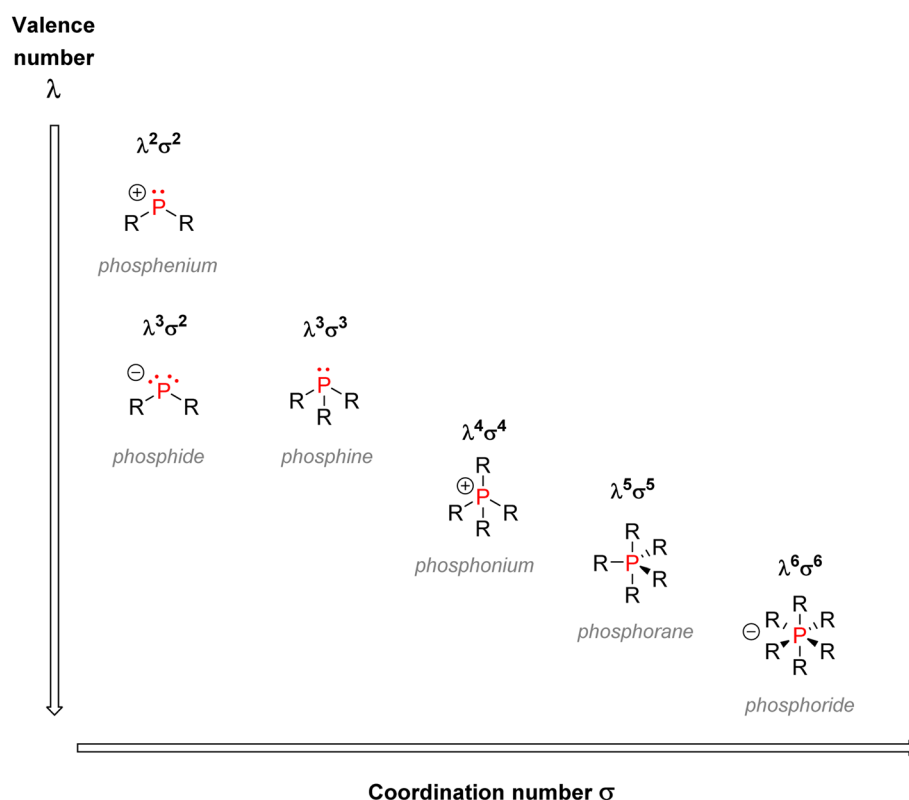


Fig. 9 Valence & coordination numbers for organophosphorus species.



less frequently used nomenclature exists. A few relevant examples will be shown according to the valence classification of organophosphorus (Fig. 9). This convention uses valency number λ and coordination number σ : thus, P(III) species are $\lambda^3\sigma^3$ and P(V) $\lambda^5\sigma^4$. The $\lambda^5\sigma^5$ and $\lambda^6\sigma^6$ have only been established in the 1960s.⁴⁶ They are still marginal but have found some recent applications as the synthesis of phosphamünchnones or montréalones developed in the Arndtsen lab or dihydridophosphoranes isolated by the Radosevich group.^{47,48} More exotic and transient structures carrying P–C double or triple bonds have been discovered such as methylene phosphines $\lambda^3\sigma^2$ or methyldyne phosphines $\lambda^3\sigma^1$.⁴⁹ Furthermore, phosphorus-containing heterocycles like aromatic phosphinine (also called phosphorine or phosphabenzene), phosphole or benzophosphole have also been well-studied (Fig. 10).⁵⁰

2.2. Trivalent phosphorus: a versatile lone pair of electrons

2.2.1. Why phosphines are excellent ligands.

As stated earlier, trivalent organophosphorus species such as PR_3 carry a lone pair which gives them nucleophilic, Lewis base, and reductive characteristics. Compared to amines, where the inversion of configuration can happen at room temperature, trivalent organophosphorus are relatively stable below 100 °C and thus can be considered a chiral center if the three substituents are different.⁵¹ Their application as nucleophiles or as reagents is well-known and beyond the scope of this review. Instead, we will focus on why P(III) compounds are extensively used in catalysis as they make excellent ligands for organometallic complexes. Phosphines, like amines, can use that lone pair to donate electrons to the metal center. The main difference resides in the fact that, unlike amines, phosphines can act as π -acids and create back bonding interactions depending on the R substituent and the σ^* orbital of the P–R bond.⁵² Electrons from the d-orbitals on the metal will then be donated to this empty σ^* orbital accordingly. The π -acidity increases according to the electronegativity of the R substituent, making PF_3 virtually as good as carbon monoxide in this aspect.

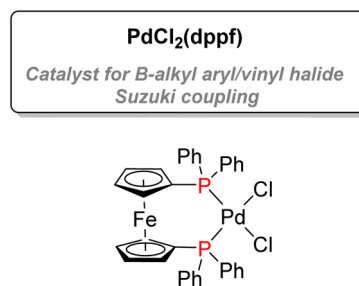
In contrast to this back-bonding effect, electron-donating substituents can enhance the σ donation of the phosphorus to the metal; this can be measured through the Tolman electronic parameter.⁵³ Changing substituents can also change geometrical configurations of the organometallic complex: bulky substituents can widen the solid angle – called Tolman angle – between the metal center at the vertex and the cone formed by the phosphine substituents.⁵⁴ As such, bulky substituents can saturate the metal's coordination sphere and

affect its catalytic properties. Chelating diphosphine ligands can also be adjusted according to their bite angle.⁵⁵ Consequently, phosphine ligands are very popular for catalyst design: they convey facile tunability of electronic and steric parameters thanks to their highly systematic and predictive behavior when playing with the substituents.

2.2.2. Ligand effects in notable catalytic systems.

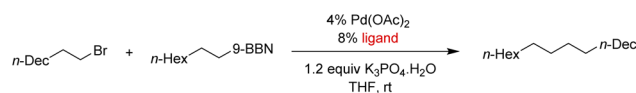
A closer look at two representative examples can illustrate the utility of phosphines as ligands. In the case of the $\text{Csp}^3\text{-Csp}^2$ Suzuki cross-coupling, $\text{PdCl}_2(\text{dppf})$ was found to be an efficient catalyst between boron-alkyl derivatives and vinyl or aryl halides/triflates.⁵⁶ The bidentate dppf ligand used can favor reductive elimination vs. a competitive β -hydride elimination thanks to its large bite angle by enforcing a cis geometry between the alkyl and vinyl/aryl substituent and by reducing the angle between them (Fig. 11).⁵⁷

Diphosphine ligands were found to be ineffective in the case of $\text{Csp}^3\text{-Csp}^3$ couplings. Two factors hinder this reaction: slow oxidative addition of the alkyl halide to the palladium center and a quick β -hydride elimination can occur instead of the expected transmetalation step. In this case, bulky electron-rich phosphines such as PCy_3 or $\text{P}(i\text{-Pr})_3$ were found to promote the cross-coupling product (Scheme 5).⁵⁸ A better σ -donation of the phosphine increases the electron density at the palladium center and enhances the oxidative addition step by promoting



dppf = 1,1'-Bis(diphenylphosphino)ferrocene

Fig. 11 Palladium–diphosphine complex used for Suzuki coupling.



Ligand	solid angle θ	CO v, cm^{-1}	% GC Yield	<i>n</i> -Dec-1-ene
BINAP	/	/	<2	<2
dppf	/	/	<2	12
$\text{P}(\text{OPh})_3$	128	2085	<2	<2
$\text{P}(n\text{-Bu})_3$	132	2060	9	27
PPh_3	145	2069	<2	<2
$\text{P}(i\text{-Pr})_3$	160	2059	68	6
PCy_3	170	2056	85	<2
$\text{P}(t\text{Bu})_3$	182	2056	<2	21
$\text{P}(o\text{-tol})_3$	194	2066	<2	14

Scheme 5 The electronic and steric effects of phosphine ligands on $\text{Csp}^3\text{-Csp}^3$ Suzuki coupling.

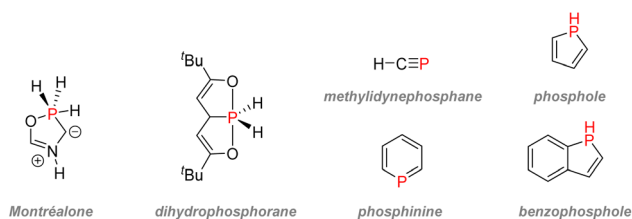


Fig. 10 Examples of phosphorus-containing heterocycles.



the formation of a coordinatively and electronically unsaturated complex. The steric hindrance of bulky phosphine also disfavors β -hydride elimination olefin byproduct by blocking empty coordination sites.

In the case of Noyori's system for asymmetric hydrogenation of diketones with BINAP/Ru, bulky and chiral diphosphine BINAP have been used to promote enantioselective transformations (Scheme 6).⁵⁹ BINAP–ruthenium(II) dicarboxylate complexes were isolated and analyzed by X-ray crystallography, showing that the rigid configuration of the ligand backbone completely blocks off any coordination on one face on the complexes.⁶⁰ It only allowed for coordination on two quadrants of the remaining face due to the steric hindrance produced by phenyl substituents.⁶¹ Thus, the use of chiral phosphines can favor topic discrimination for substrate coordination in a process called dynamic kinetic resolution. Those stereoselective processes can yield highly enantiopure products. This technique was used to synthesize L-DOPA by Knowles for the industrial synthesis of (–)-menthol.^{62,63}

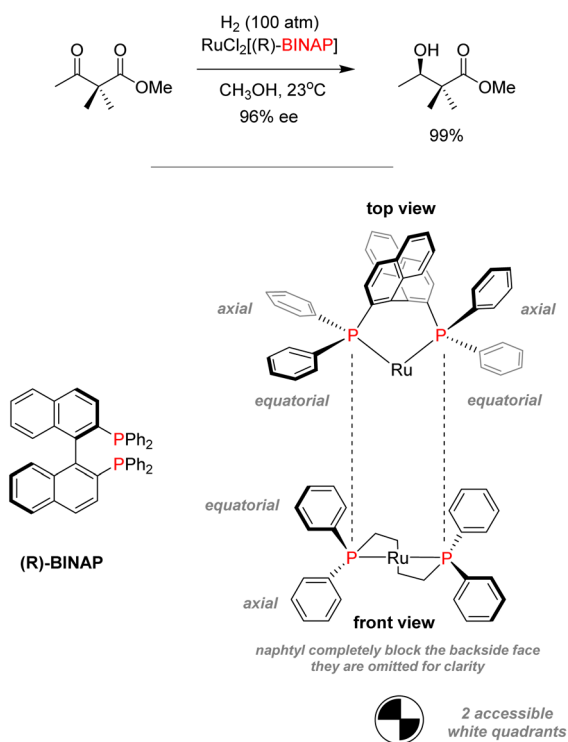
In addition to classical diphosphine ligands, another noteworthy class of P(III) ligands discovered in the 1970s has proven particularly interesting and has been the subject of many recent advances in designing efficient catalytic systems.^{64–68} Pincer ligands are tridentate ligands which coordinate metal centers tightly with a meridional configuration, thus achieving better thermal and air stability. Moreover, the restricted coordination sites of such organometallic complexes often limit the generation of unwanted side products. For P(III) containing pincer ligands, the central chelating atom can either be a carbon for

PCP ligands or a nitrogen for PNP ligands. Akin to their simpler diphosphine counterparts, pincer ligands can easily be tuned by selecting appropriate substituents on the phosphorus for electronics & sterics or by designing the central framework and spacers for specific geometries or electronics. Such ligands activated relatively unreactive bonds such as C–C, H–H or C–F bonds. Milstein's direct dehydrogenative synthesis of amides from alcohol and amines in 2007 was a significant milestone.⁶⁹ Those early discoveries paved the way for the development of various applications using organometallic pincer complexes for more sustainable catalysis: among those are atom-economic hydrogen transfer reactions such as hydrogenation & dehydrogenation of organic molecules, efficient cross-couplings, hydrogen storage materials, greenhouses gas conversion, or plastic depolymerization.^{70–72} Novel development, including the use of non-symmetrical frameworks or merging metallocenes with pincer ligands has also been reported.^{73,74}

2.2.3. Other applications. A few other niche applications in material chemistry of heterocyclic P(III) species, such as phospholes, will be detailed in a further section. It is also worth mentioning that bulky P(III) compounds are often combined with sterically hindered Lewis acids to form Frustrated Lewis Pairs (FLP). Those FLPs can split H₂ and activate other molecules; as such, they have found applications as metal-free catalysts for hydrogenations, hydrosilylations, or CO₂ capture.^{75,76}

2.3. Pentavalent phosphorus: a stable phosphoryl bond

2.3.1. Properties of the phosphoryl bond. While P(V) molecules can be used as ligands akin to P(III) ones,^{77,78} their applications are more diverse than coordination chemistry and catalysis. Pentavalent organophosphorus have in common the presence of phosphoryl P=O bond, and more occasionally, other chalcogens such as P=S or P=Se. Contrary to amine oxides which necessitate specific oxidizing agents to be formed but can easily be reduced, trivalent phosphines can readily be oxidized under air (or by almost any other oxidant) to form P=O bearing phosphine oxides. However, their reduction reveals much more demanding and requires very strong reductants such as silicon hydrides or lithium aluminium hydrides. Phosphoryl bonds are considered as mostly stable and somewhat unreactive. In that aspect, they are not similar in reactivity with C=N or C=O bonds as P=O does not undergo nucleophilic addition easily. However, oxygen can react with strong acids or potent electrophiles. The stability comes from a high bond dissociation energy, varying between 129 to 139 kcal mol⁻¹, which can explain the oxophilicity of trivalent phosphorus and their tendency to form P(V) adducts.⁷⁹ The formation of P=O bonds is usually a decisive factor for many reactions, e.g., the driving force of the Wittig reaction or the Wittig–Horner reaction. The stability of the phosphoryl bond was somewhat of a controversial topic for past chemists: early literature drew P=O bond as shown (Fig. 12a), then a later convention was to describe phosphoryl bonds as undergoing a back bonding of sort between a pair of electrons from the oxygen and the empty 3d orbitals on phosphorus as shown by



Scheme 6 Noyori asymmetric hydrogenation of diketones with ruthenium–BINAP complexes.



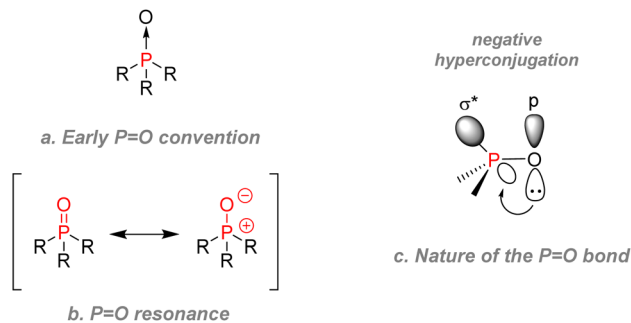


Fig. 12 Drawing conventions and nature of phosphoryl P=O bond. (a) Early convention (b) resonance (c) nature of the P=O bond.

the resonance equation (Fig. 12b). To this date, many phosphorus specialists argue that the zwitterionic form describes phosphoryl bonds the best.^{80,81}

However, computational chemistry showed that hyperconjugation was probably between the oxygen lone pair electrons and the empty anti-bonding σ^* orbitals of the P–R bonds (Fig. 12c).^{82,83} Overcoming the controversy are several unequivocal metrics: phosphoryl bonds are also short: 1.475 to 1.490 Å compared to 1.60 Å for a typical single P–O bond from a P–O–P unit. Regarding their dielectric constant, phosphoryl bonds are also highly polar (4.51 D in triphenylphosphine oxide).⁷⁹ The phosphoryl oxygen is thus highly susceptible to forming hydrogen bonds, which can make P(v) species prone to water solubility in some cases. Those characteristics make the ³¹P NMR of the phosphoryl group shift according to the solvent as it is sensitive to protic environments.⁸⁴ Finally, we can classify phosphoryl moieties in terms of substituent effects: the inductive donating effect of the phosphinyl group (where R = C for R₂P(O)–) is comparable with the methoxy group for a saturated system.

Nevertheless, phosphinyl groups are moderately electron-withdrawing by resonance, making them *meta* director and deactivator in the case of electrophilic aromatic substitutions.^{85,86} Alternatively, the strong electron-withdrawing effect of phosphoryl bonds enables them to stabilize carbanions or saturated carbons at the α position. Those properties make the phosphoryl bond a singular one in nature and will explain the bioactivity of many phosphorylated molecules. The following sections will detail several applications of those P(v) containing structures.

2.3.2. Medicinal chemistry. Despite the relatively recent use of P(v) organophosphorus in chemotherapeutics, they have succeeded as antitumor reagents, antibiotics, bone resorption inhibitors, antimalarial agents, antidepressants, and antihypertensive agents, and antivirals. Several factors can explain the interest of medicinal chemists in pentavalent phosphorus (Fig. 13). The first one is using P(v) such as phosphonates as analogs of biologically abundant phosphates. Because of their similar size and bond angle, a CH₂ group bound to the phosphonate is isosteric to the oxygen bound to phosphonates, making both structures virtually superimposable. However, the C–P bond from phosphonates is much more stable to hydrolysis

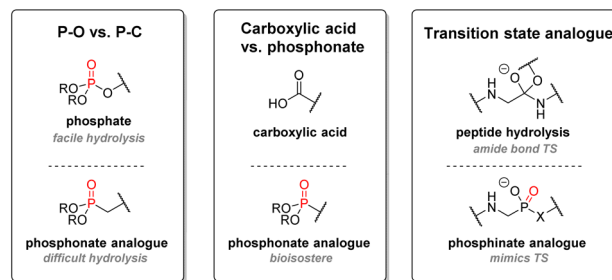


Fig. 13 Strategies for P(v) organophosphorus use in medicinal chemistry.

than P–O, making phosphonates great candidates as phosphates mimics and antagonists. Another P(v) structure of interest is phosphonic acids: they can also be considered bioisostere of carboxylic acid and thus can inhibit biochemical processes relying on those acids. Therefore α -amino phosphonates are considered as bioisosteres of amino acids. Finally, tetrahedral P(v) structures also excel at being transition state analogs. By mimicking the transition state of a substrate, such as the Csp³ intermediate of a hydrolyzed peptide bond, and because P–C or P–O bonds are longer than their corresponding C–C and C–O counterparts, a rationally designed therapeutic can fit into the active site and stay there, thus inhibiting enzymatic processes.^{79,87} Phosphonates, phosphinates, and phosphoramidates are effective for this strategy. Overall, most phosphorus-containing pharmaceuticals are phosphates, phosphoramides, or phosphonates, whereas phosphinates or phosphine oxides remain rare.⁸⁸

In terms of pharmacokinetics, P(v) structures exhibit high polarity and thus generally good aqueous solubility. They have also been found to have a sufficiently long half-life in human liver microsome conditions, allowing for good bioavailability before metabolism. With a few exceptions, most structures containing P(v) moieties were also usually stable under acidic or basic conditions. Functional groups containing acidic phosphorus such as phosphates, phosphinates, or phosphonates suffer from the significant drawback of turning into charged species under biological conditions, thus decreasing cell permeability and oral bioavailability. One way to circumvent those inconveniences is to use solubility-augmented prodrugs, which will be metabolized into the active therapeutic compound. Phosphine oxide-based drugs are still infrequently used, but recent examples have shown promising results.⁸⁸

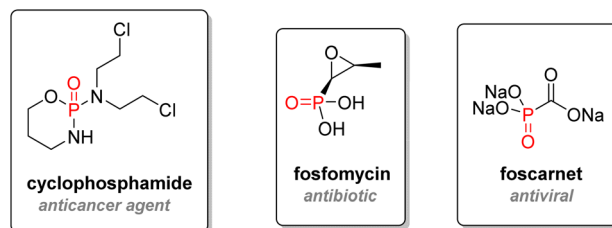


Fig. 14 Early organophosphorus therapeutics.



Tutorial Review

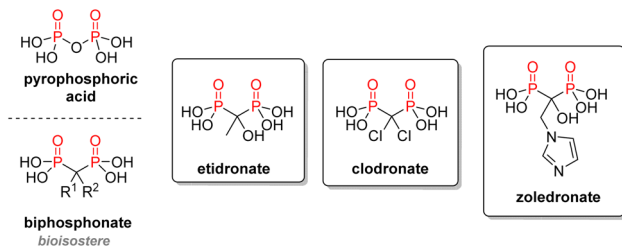


Fig. 15 Bisphosphonates as bone resorption inhibitors.

2.3.2.1. Early hits. The cyclic phosphorodiamidate anticancer drug called ‘cyclophosphamide’, discovered in 1958, is used as an immunosuppressor drug to treat cancers, leukemia, and lymphoma (Fig. 14). After metabolism, cyclophosphamide is converted to acrolein and an active chlorinated phosphorodiamidate, which acts as a nitrogen mustard compound.⁸⁹ Cross-linking with DNA strands will occur, and cell death will ensue. Cyclophosphamide figures in the World Health Organization’s list of essential medicines.⁹⁰ Another entry in this list is Fosfomycin, an epoxide phosphonic acid used as an antibiotic to treat bladder infections.⁹¹ Fosfarnet, a very simple trisodium phosphonate salt first synthesized in 1924, was found in 1978 to be an effective antiviral agent by inhibiting viral DNA polymerase.⁹² It is effective against herpes and has been used for salvage therapy treatments against HIV.⁹³

2.3.2.2. Bisphosphonates. Following the rationale of substituting oxygen by carbon between two P(v) to mimic phosphates, the idea of using bisphosphonic acid derivatives as analogs of pyrophosphoric acid emerged. Similar to pyrophosphoric acid, bisphosphonates exhibit an intrinsic affinity for calcium ions and accumulate very well in bones, where they inhibit excessive bone resorption (Fig. 15). This class of drugs is used against osteoporosis, Paget’s disease of bones, or other bone disorders.⁹⁴ The first generation of bisphosphonates included derivatives of etidronic or clodronic acids bearing simple R¹ and R² substituents on the central carbon. Over the years, a better structure–activity relationship was established, and the current drugs have a dramatically increased potency, going up to a 10 000-fold upgrade compared with etidronate in the case of zoledronate.⁹⁵ The current strategy is to use a hydroxyl R¹ substituent to better the affinity with calcium or magnesium ions, while using nitrogen in the R² substituent will increase the hydrogen bonds with the target enzyme. Bisphosphonate moieties have also been cross-linked to other drugs to take advantage of the exceptional binding affinity of these structures to the hydroxyapatite from bones. This method was coined osteoporotic drug delivery system.⁹⁶

2.3.2.3. Aminophosphonates. Another crucial structural motif of organophosphorus drugs is aminophosphonates: α -aminophosphonic acids are bioisosteres of α -amino acids and resemble the high-energy transition state of the hydrolysis of ester and amide bonds. The simple amino-acid mimetics strategy is advantageous in the case of antibacterial drugs (Fig. 16). For instance, alafosfalin, which is a phosphonodipeptide, releases L-aminoethanephosphonic acid after metabolism.

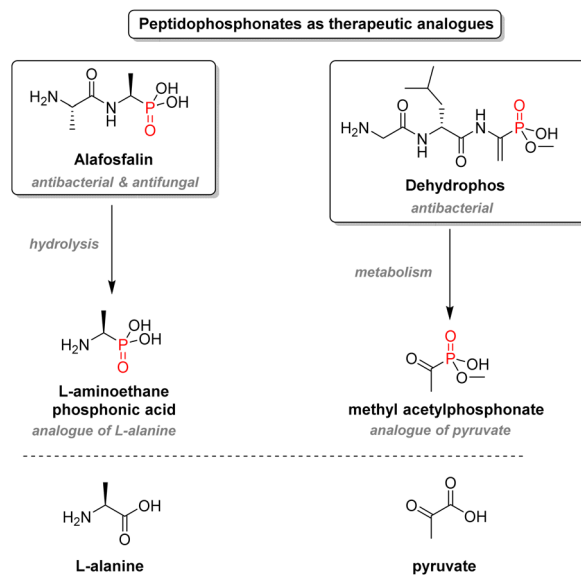


Fig. 16 Peptidophosphonates as therapeutic analogs.



Scheme 7 Metabolism of Fosinopril.

The phosphonate behaves like the natural substrate for L-alanine racemase except for its higher binding affinity to the cofactor, thus inhibiting the enzyme and subsequent cell wall biosynthesis.⁹⁷ Dehydrophos, a peptide analog produced by *Streptomyces luridus*, also shows wide-spectrum antibacterial activity.⁹⁸

Fosinopril is a successful β -aminophosphinate therapeutic used to treat hypertension and is commercialized by Bristol-Myers Squibb. The compound is administered as a prodrug: the phosphonic acid is esterified in order to improve bioavailability, and subsequent metabolism will cleave this P–O bond to release the active compound, Fosinoprilat (Scheme 7). This molecule is designed to be an analog of the transition state emerging from the hydrolysis of the dipeptidic substrate from the Angiotensin-Converting Enzyme (ACE).⁹⁹



Fig. 17 Incadronate and lexidronam.





Fig. 18 Nucleotide analogue antivirals.

Another mechanism of action for organophosphorus drugs is to be used as an irreversible inhibitor. This strategy includes the most nefarious examples of chemical warfare: nerve agents blocking the acetylcholinesterase enzyme. Details about their

mode of action will be explained below in the agrochemistry application section. As for therapeutics, diaryl α -amino-phosphonates exhibited promising inhibition of serine protease enzymes, which could lead to several applications such as anti-cancer agents, therapeutics against inflammatory disease, infections, or against hypertension.^{100–102} Finally, amino-phosphonates can also be used for treatment against bone disease: incadronate merges both aminophosphonate and bisphosphonate structures and has been used against hypercalcemia and cancer development *in vitro* (Fig. 17).^{103,104} Macrocyclic polyaminophosphonates chelating ligands have been used with heavy lanthanide nuclei for either imagery or targeted radiotherapy of bone cancer.^{102,105,106}

2.3.2.4. Antiviral nucleotide analogs. Organophosphorus nucleotide analogs are also a prevalent drug design strategy for antivirals. Those analogs behave and compete with natural nucleotides (or deoxynucleotides) to be incorporated into either viral DNA or RNA strands (Fig. 18). After phosphorylation steps, the creation of the consecutive phosphodiester bond with a new nucleotide for the strand is inhibited, either because of a missing 3'-hydroxyl group on the drug or because of steric interactions on the carbohydrate mimic. This mode of action thus blocks the growth of the viral strand and terminates the chain. This approach successfully inhibited viral reverse transcriptase in the case of hepatitis B: adefovir dipivoxil and tenofovir disoproxil are both phosphorylated adenosine derivatives prodrugs, and cleavage of the pivaloyloxymethylated P-O bonds will occur through metabolism, thus releasing the active forms of the drugs.¹⁰⁷ Tenofovir can also be used against HIV. Remdesivir is a phosphoramidate prodrug synthesized by Gilead, which metabolite mimics adenosine triphosphate to interfere and block viral RNA polymerase. It was initially



Fig. 19 Mode of action of organophosphorus acetylcholinesterase inhibitors.



designed as a treatment for hepatitis C or Ebola virus but was later tested against SARS-CoV-2.^{108,109} Similar mechanisms of action are at work with other FDA approved phosphoramidate nucleotide drugs such as Sofosbuvir (blocking RNA polymerase from hepatitis C) or Cidofovir (blocking DNA polymerase from cytomegalovirus).^{110,111}

2.3.3. Agrochemicals

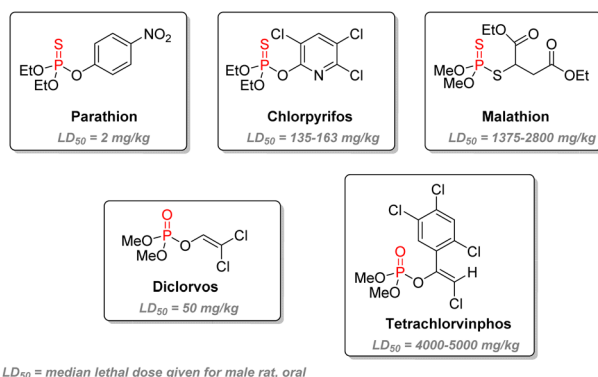
2.3.3.1. Insecticides. The usage of organophosphorus species – often generally coined “organophosphates” – for agrochemical purposes dates back to WWII. As previously stated in Section 1.1.2, Schrader in Germany and Ghosh in the UK independently investigated the use of phosphoryl derivatives as insecticides before the war and discovered that the mammalian toxicity of those compounds was surprisingly exceptionally high.¹¹² Tabun, a clear colorless liquid phosphoramidite, was the first organophosphorus chemical intended as an insecticide which was repurposed as a warfare nerve agent. Regrettably, it paved the way for developing the whole G-series and V-series of lethal organophosphorus chemical weapons. However, proper substituent modification on the organophosphorus structure could reduce mammalian toxicity, thus allowing for a more specific and selective pesticide application.⁷⁹ All organophosphorus insecticides act as irreversible acetylcholinesterase inhibitors: they all mimic the molecular shape of the neurotransmitter acetylcholine and can dock in the enzymatic active site (Fig. 19). However, instead of having the serine residue hydrolyzing acetylcholine, serine gets phosphorylated, thus impeding the usual regulation of acetylcholine and finally causing muscle paralysis and eventual death of the insect.¹¹³

As electrophilic phosphorylating agents, most insecticides of the organophosphorus family are either phosphorothioate (one sulfur replacing one oxygen atom: P=S replacing P=O, or P-S replacing P-O), phosphorodithioate, or vinyl phosphates (Fig. 20). Three key features play a role in the design: firstly, usually, two methyl, ethyl esters, or amino residues are present in order to balance the solubility, as free hydroxyl or thiol group would be ionized under physiological conditions, thus impeding a good bioavailability. Substituents also determine the rate at which the phosphorylated acetylcholinesterase

serine residue goes back to its active state: hours for methyl esters to days with ethyl esters.¹¹⁴ Secondly, the phosphoryl group can also be replaced by P=S. As electrophilicity of the P center is crucial for bioactivity, P=S compounds are less reactive towards acetylcholinesterase. However, it will take several hours for *in vivo* oxidase to convert thions P=S to active oxons P=O compounds.¹¹⁴ This delay allows for lesser mammalian toxicity as part of the insecticide can be eliminated during this process while still being active against faster-metabolizing insects. Thirdly, the remaining substituent should act as an effective leaving group which can be substituted by serine. Efficiency is increased with electron-withdrawing groups. This explains why the *p*-nitrophenoxy substituent of parathion makes it highly active and toxic, resulting in its global ban from sales and imports. The emergence of malathion in the 1950s is an important landmark as it bears significantly lower toxicity towards humans. It is suitable for general purposes such as household or garden use but has also been used for public health in the struggle against mosquitoes or flies. Chlorpyrifos, developed in the 1960s, is still currently one of the most widely used agricultural pest control agents and the first organophosphate used in the US in 2012,¹¹⁵ despite increasing studies highlighting long-term exposure health problems. Another subcategory of P(v) insecticides are the vinylphosphates, including the highly volatile and regulated dichlorvos and the less toxic tetrachlorvinphos.

Organophosphorus insecticides react with air, light, and water and are more easily hydrolyzed and biodegraded than organochlorine insecticides, which significantly impact the environment. For instance, formerly extensively used but now banned organochlorine insecticide DDT has a 2.8 year half-life in soil vs. 1 or 2 weeks for malathion or 3–6 months for chlorpyrifos.¹¹⁶ However, their substantial use has shown contaminations in grains, vegetables, fruits, and milk. The ongoing concern with organophosphorus insecticides is their acute human toxicity: yearly worldwide, organophosphate poisoning has been estimated to be 3 million, causing 200 000 deaths.^{117,118} Detrimental effects on bees, fish, birds, and wildlife have also been reported.¹¹⁹

2.3.3.2. Herbicides. Another important application for P(v) compounds in agricultural chemistry is herbicide (Fig. 21). The most successful herbicide is glyphosate – commercially sold as Roundup – an α -aminophosphonate type systemic herbicide first marketed in 1974. It is not only the world's currently most used herbicide but also pesticide overall by weight. It has been



LD_{50} = median lethal dose given for male rat, oral

Fig. 20 Examples of organophosphorus insecticides and their corresponding LD_{50} .

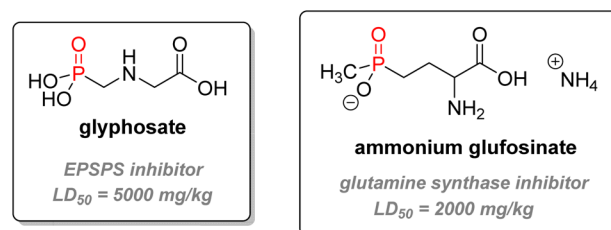


Fig. 21 Common organophosphorus herbicides.



the most used herbicide in the US for agriculture since 2001, for the industry/commercial/government market in 2009–2012, and it was ranked second for home and garden market in the same period.¹¹⁵ The mechanism of action of the phosphonate is to inhibit 5-enolpyruvyl-shikimate-3-phosphate synthase (EPSPS), an enzyme part of the amino-acid biosynthetic pathway from glucose. This enzyme is specific to plants and does not exist in mammals. The efficient disruption of the amino-acid synthesis causes the death of many plants without selectivity, making glyphosate a powerful broad-spectrum herbicide.¹²⁰ Thousands of compounds have been tested to inhibit EPSPS, but only glyphosate has been commercialized so far. Transgenic glyphosate-resistant crops such as soybean, maize, cotton, or canola breeds are commonly used with the potent herbicide.¹²¹ Key features include good uptake through the leaves, excellent translocation (movement of the herbicide inside the plant from the uptake location to the site of action), the absence of natural resistance amongst higher-plants, very low toxicity (LD₅₀ oral for male rats above 5000 mg kg⁻¹),¹²² and a relatively short half-life in soil (30–174 days).¹²³ However, recent controversies have emerged related to the possible human carcinogenicity and environmental issues linked to the use of glyphosate.^{124,125} Contradictory studies and changing legal status have shown the need to reassess data about the herbicide standards.^{126,127}

Glufosinate is a non-selective phosphinate herbicidal compound of lesser importance, which can also be used as an antibiotic or a fungicide. It is the active metabolite of bialaphos, a natural peptide derivative synthesized by *Streptomyces hygroscopicus*.¹²⁸ It acts as a glutamine synthase inhibitor, impeding photosynthesis and causing the death of the plant. Glufosinate ammonium bears a lower LD₅₀ than glyphosate (416–2000 mg kg⁻¹ for rat, oral).¹²⁰ Other organophosphorus compounds used as agrochemicals include fungicide Fosetyl-Al or plant growth regulators ethephon or glyphosine (Fig. 22).

2.3.4. Organic electronics & materials with optical properties. The use of organophosphorus for organic electronics is a recent one and is still underdeveloped. A few examples of this section technically belong to the P(III) category, but they have been summarized here along with the P(V) species for a more consistent thematic clarity. The design of organic electronic materials necessitates the synthesis and fine-tuning of organic structures possessing desirable electronic properties such as conductivity or light emission. In addition, they are potentially much cheaper than inorganic electronics, and printed electronics can also be flexible. This part will focus on phosphine

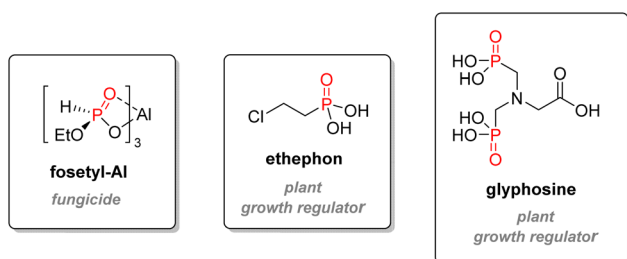


Fig. 22 Examples of organophosphorus agrochemicals.

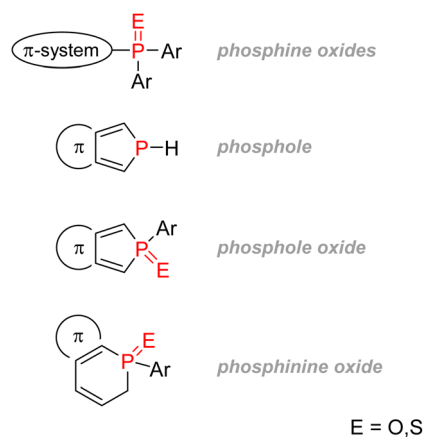


Fig. 23 Common organophosphorus structures for organic electronics.

oxides, phospholes, phosphole oxides, and phosphinine oxides used as organic electronics (Fig. 23).

2.3.4.1. Organic light-emitting diodes. Currently used organic light-emitting diode (OLED) technology has been a significant advance in display and light-source technology. Electrons injected from the cathode will be transported through an electron transport layer to the emissive layer. Similarly, the lack of electrons, known as holes, will be injected by the anode and transported through the hole transport layer to the emissive layer as well. This emissive layer will recombine both electrons and holes into excitons, radiating light according to the dye used in the layer (Fig. 24). However, conventional OLEDs rely on light emitted by fluorescence, limiting the conversion of excitons into light only to the singlet excitons, representing a 25% internal quantum efficiency. Another improved strategy exists with phosphorescent organic light-emitting diodes (PhOLED), which can harness the remaining triplet state excitons and increase efficiency. Nevertheless, phosphorescent dyes (called guests or dopants) inside the emissive layer should be spaced out inside a matrix (host) in order to avoid triplet-triplet quenching between the excited molecules.¹²⁹

2.3.4.1.1. Phosphine oxides. Phosphine oxides have been used both in the electron transport layer and in the host

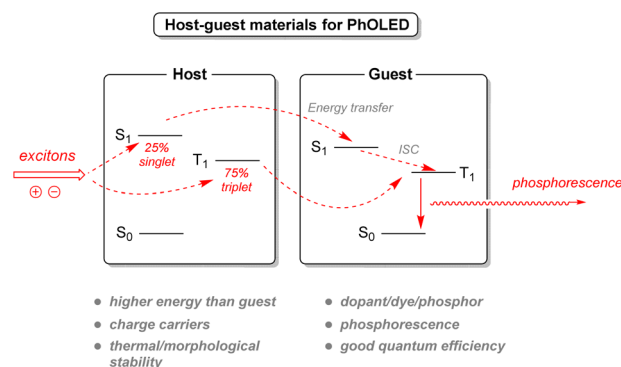


Fig. 24 Roles of host-guest materials for PhOLED.



material: the high polarity and inductively withdrawing effect of the P=O bond makes phosphine oxide especially well-suited for electron transport. Their high triplet state energy can also effectively capture triplet state excitons and redistribute them to the lower energy triplet state of the light-emitting guest molecule through energy transfer. Thanks to their tetrahedral geometry, phosphine oxide can interrupt conjugation and give good thermal and morphological stability with their steric bulk while keeping a necessary high triplet state energy. In order to improve charge transport, some host materials can be 'bipolar', and transport both electrons and holes by using both an electron-withdrawing group such as a phosphine oxide moiety and an electron-donating one such as a tertiary amine: for instance (Fig. 25(1)) combines an active fluorophore bifluorene backbone with a diphenylphosphine oxide. It has been used as an electron transport host material for high-efficiency blue electrophosphorescent devices.¹³⁰ Bipolar host materials using hole-transport moieties such as carbazole and electron transport ones like phosphine oxide have been used to construct deep blue colored PhOLED (Fig. 25(2)).¹³¹ Polymeric bipolar hosts combining carbazoles and phosphine oxides have also shown good charge transport capabilities and a high triplet state (Fig. 25(3)).¹³² Phosphine oxides have also directly been used as light-emitting material (Fig. 25(4)) by fusing them with fluorene or carbazole in the architecture for blue emission,^{133,134} or by directly incorporating classic guest phosphorescent complexes such as blue FIrPic and yellow [(fbi)₂Ir(acac)] inside a fluorinated phosphine oxide polymeric backbone to give an efficient single polymer for white PhOLED.¹³⁵

The excellent review from Jeon & Lee compiled many phosphine oxide-containing materials used for PhOLEDs and their physical data.¹³⁶ Despite the outstanding abilities of phosphine oxide-based PhOLED material in terms of high triplet state energy and electron transport, their stability towards UV or electrical stress is compromised when P-C bond cleavage occurs during excited or charged states.^{137,138} This could significantly lower the life-time of blue PhOLED and remains an ongoing obstacle to be overcome.

2.3.4.1.2. Phospholes. Phosphorus-containing heterocycles have also found applications in organic electronics. Phosphole,

the phosphorus relative of pyrrole and thiophene, is a P(III) 5-membered ring bearing very weak aromaticity: 1*H*-phosphole has a calculated nucleus-independent chemical shift (NICS)(0) of -5.6 compared with -11.9 with furan or -13.3 for thiophene, which is more aromatic. The reason for such low aromaticity for phosphole and phosphole oxides is that P(III) has a high energy inversion barrier and keeps a pyramidal configuration, disrupting a good overlap of the π electrons.¹³⁹ Contrary to pyrrole, the phosphorus lone-pair does not participate in the aromaticity, as it is instead an interaction between the diene π electrons and the σ^* from the P-R bond that takes place (Fig. 26). Changing the P-R exocyclic substituent can change this very $\sigma^*-\pi$ interaction and thus tune the aromaticity of the phosphole ring.

Moreover, phospholes can act as standard platform structures for obtaining diverse modified π -conjugated materials as functionalization on the phosphorus lone pair is easy: oxidation (oxygen, sulfur, or selenium), alkylation, or coordination with a transition metal are all possible. Finally, due to low LUMO energy, phospholes have electron acceptor properties that can also be modified according to the phosphorus's P-R substituents or oxidation state. Those properties make phospholes and phosphole oxides containing π -systems versatile tools for material chemists.

Both dibenzooxophosphole, used as an electron acceptor unit, and carbazole, used as a hole acceptor unit, were integrated into bipolar host material for OLED purposes with good efficiency in combination with Os(II) guest dyes (Fig. 27(1)).¹⁴⁰ Highly performant thioxophosphole (Fig. 27(2)) was directly used as a fluorescent dye for OLED and showed the highest quantum efficiency for phosphole emitting layer material.¹⁴¹

2.3.4.2. Dye-sensitized solar cells, bioimaging, electrochromic device, and other stimuli-responsive materials

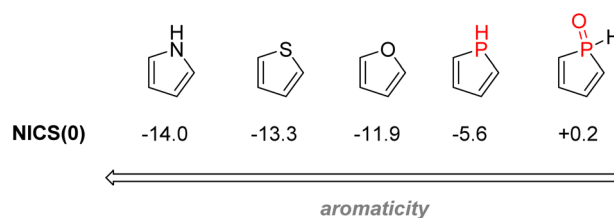


Fig. 26 Calculated aromaticity of phospholes and phosphole oxides compared to other electron-rich heteroarenes.

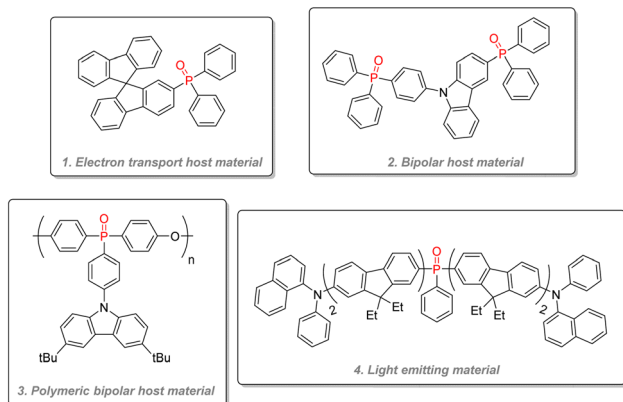


Fig. 25 Examples of phosphine oxides as PhOLED materials.

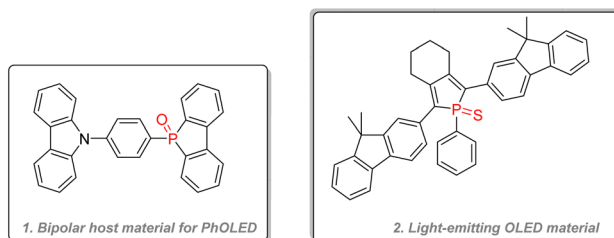


Fig. 27 Phospholes as OLED materials.



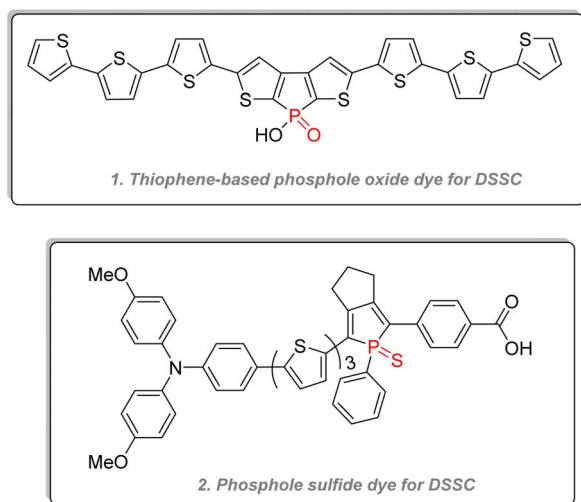


Fig. 28 Phospholes used for dye-sensitized solar cells.

2.3.4.2.1. Phospholes. Aside from OLEDs, phosphole oxides and sulfides have also been used as dyes for dye-sensitized solar cells (DSSC) (Fig. 28),^{142,143} for electrochemical cells, and electrochromic devices.^{144,145} Polyaromatic phospholes and phosphole oxides have also been used in reversible stimuli-responsive color-changing structures, which can be indispensable for smart materials such as display, thermochromic paints, or self-dimming glass. Reus & Baumgartner detailed many examples of those stimuli-responsive chromic organophosphorus architectures in their review, showing among other interesting properties electrochromism (change of color after a redox reaction), photochromism (change of color according to the UV/visible light irradiation), thermochromism (change of color according to the temperature), halochromism (change of color according to the pH) or even mechanochromism (change of color according to the mechanical stress).¹⁴⁶

2.3.4.2.2. Phosphinine derivatives and P-containing 6 membered rings. Phosphorus-containing six-membered rings such as phosphinine and phosphinine oxide bear unique optoelectronic properties. Jiang reported in 2015 the use of a phosphorus dipyrromethene and its aza counterpart as fluorescent marker dyes for bioimaging (Fig. 29(1)). The compound showed better water solubility than the currently used BODIPY dye.¹⁴⁷ Other phosphorus versions of traditional dyes have been synthesized, such as phospho-rhodamine or phospho-fluorescein, and all have shown good quantum efficiency, water solubility, and excellent resistance towards photo-bleaching.^{148,149} A very ingenious strategy to detect calcium ions in living cells was to combine a far-red fluorescent phospho-fluorescein scaffold with calcium probes (Fig. 29(2)).¹⁵⁰ Polyaromatic systems incorporating phosphinine derivatives have recently been studied by Romero-Nieto *et al.*, and the compound (Fig. 29(3)) showed outstanding photochemical properties and good suitability for photoelectrochemical cells.¹⁵¹ An excellent review written by Regulska & Romero-Nieto focuses on the applications of extended π systems based on phosphorus-containing six-membered heterocycles.¹⁵²

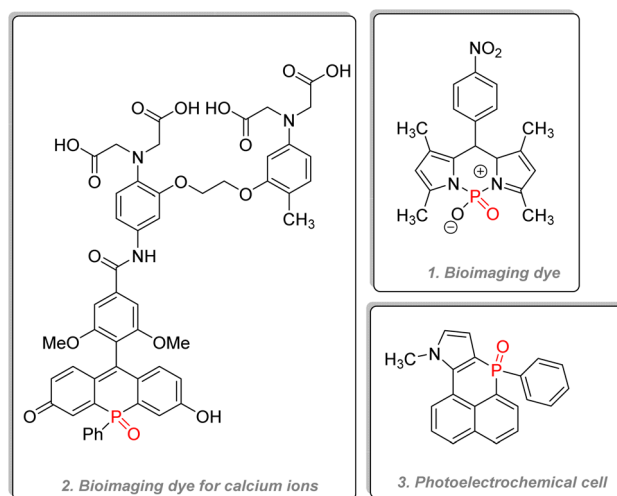


Fig. 29 Phosphinines in organic electronics.

2.3.5. Flame retardant. Flame retardants are chemicals added to materials such as polymers, textiles, or wood to delay their flammability in case of fire. They are often used in insulation foams for construction, electrics, electronics, cables, furniture, *etc.* Around 2.26 million tonnes of flame retardants were synthesized worldwide in 2018.¹⁵³ A good flame retardant must not inhibit the inherent material properties too much; it should be compatible with its processing methods while still satisfying the various safety, health, and environmental regulations. Flame retardants have different modes of action targeting either factor of the fire triangle: heat, oxidant, or combustible. A few common strategies are: reducing the overall quantity of combustible by diluting it with fillers, producing endothermic reactions, generating non-combustible gas, forming gas-phase quenching radicals to prevent oxidation, or forming a layer of stable char which protects the combustible from further combustion.

Aluminium hydroxide is a mass-produced, and inexpensive flame retardant used widely along with other metal hydrides for cheap commodity materials such as thermoplastic polyolefins. However, $\text{Al}(\text{OH})_3$ is often required in very high loadings, which can negatively affect the material properties. More advanced and sensitive materials such as engineering polymers or high-performance ones often necessitate more suitable retardants. The oldest technology used for those materials was based on halogenated flame retardants, once considered the workhorse of this field. Their decomposition into halogen acid gas acts as a radical quencher, thus obstructing the combustion process.¹⁵⁴ Despite claims from manufacturers, some popular halogenated flame retardants have proven to have a high bioaccumulation rate and to be toxic for humans and animals, causing harmful neurobehavioral effects and endocrine disruption.¹⁵⁵ This led regulatory authorities in Europe and North America to ban several brominated and chlorinated compounds, thus pushing for increased use of phosphorus-based flame retardants.

Inorganic forms of phosphorus such as red phosphorus previously used in nylon 66 or polyethylene have been used as



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flame retardants, but the hazards linked to handling and coloring have since restricted its use. Cheap and efficient ammonium phosphates, whose use as fertilizer was already covered earlier, can also be used as a flame retardant for cellulosic materials such as textile, paper, and wood. Organophosphorus flame retardants have two main mechanisms of action, one in the condensed (solid) phase and one in the gas phase. First, phosphorus will cross link between the carbon decomposition products occurring on the solid-state of the material and form inorganic glasses like polyphosphate and phosphoric acid, thus inhibiting pyrolysis. Second, in combination with other material, phosphoric acid residues can promote intumescence, insulating the unburnt parts. Volatile phosphorus fragments (PO^{\cdot} , HPO_2^{\cdot} and PO_2^{\cdot}) act as radical scavenger and quench reactive H and OH radicals to delay exothermic processes and decrease the supply of flammable gases.¹⁵⁶ Organophosphorus flame retardants can be used either as additives or as reactive.

The first category, the most used one, contains most of the classic organophosphorus flame retardants.¹⁵⁷ Triphenyl phosphate (Fig. 30(1)) is used in cellulose acetate films and engineering thermoplastics. Tricresyl phosphate is used in rubbers and flexible PVC. Resorcinol bis(diphenyl phosphate) (Fig. 30(2)) is a less volatile and highly successful compound used for engineering thermoplastics, polyamides, vinyls, thermoplastic polyesters, thermosets, and various elastomers. Bisphenol A bis(diphenyl phosphate) (Fig. 30(3)) is a cheaper option than the resorcinol diphosphate, but it requires more loading for similar properties. Phosphonates such as dimethyl methylphosphonate (Fig. 30(4)) were used in polyester resins and ductwork, but they are now more restricted. Aluminium diethyl phosphinate (Fig. 30(5)) is used in thermoplastics such as polyamides, polyesters, ABS.

Reactive flame retardants bear reactive chemical moieties, which allow them to be directly cross-linked into the material, making them an integral part of them. This often leads to a lesser impact on the physical properties of the overall material

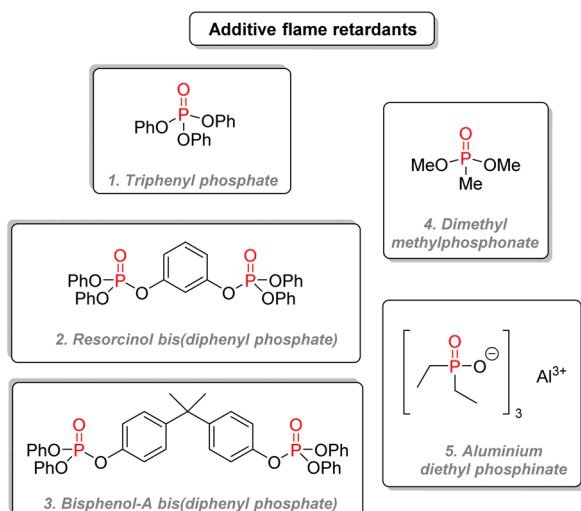


Fig. 30 P(v) organophosphorus as additive flame retardants.



Fig. 31 P(v) organophosphorus as reactive flame retardants.

and a decreased leaching into the environment. Bis(4-aminophenyl)phenylphosphine oxide (Fig. 31(1)) was used in polyurethane resins.¹⁵⁸ Phosphinate adducts such as 9,10-dihydro-9-oxa-10-phosphaphenanthrene-9-oxide (Fig. 31(2)), which is also used as an additive, can also be added as a reactive flame retardant in epoxy resins while keeping high thermal stability.¹⁵⁹ Derivatives of this structure also gave good flame retardant properties for the electrolytes of Li-ion batteries.^{160,161}

However, growing concern about the ubiquitous presence of organophosphorus flame retardants residues in outdoor and indoor environments, in addition to the rising issue of their potential toxicity, has led reports to question whether the substitution of halogenated compounds by P-flame retardants additives was the right choice.¹⁶² One interesting direction to solve this issue would be to use biocompatible and bio-resourced phosphorylated biopolymers such as casein or DNA as flame retardant coating, as a few recent reports showed.^{163–165}

2.3.6. Plasticizer. Additionally to their flame retarding properties, P(v) compounds are also used as plasticizer additives in different materials. The role of plasticizers is to confer better flexibility, workability, or distensibility and decrease viscosity or glass transition temperature. They are chosen according to various criteria such as their compatibility with the host material, their properties (mechanical, thermal, electrical, or rheological), toxicity, or costs.¹⁶⁶ Most phosphorus-based plasticizers used in the industry are phosphates: aryl phosphates, alkylaryl phosphates and alkyl phosphates.¹⁶⁷ Aryl phosphates combine both flame retardant properties with plasticizing ones. Aryl phosphates such as triphenyl phosphate are used in hydraulic fluids, lacquers, varnishes, nail polish, and tricresyl phosphate is used for PVC and rubber. It is worth noting that the *ortho*-cresyl version is highly toxic and has mostly been banned for use.¹⁶⁸ Alkyl arylphosphates exhibit better abilities as plasticizers but are poorer flame retardants. 2-Ethylhexyl diphenylphosphate is the most commonly used example of this category; it is also used in flexible PVC,



Fig. 32 Organophosphorus as extraction ligands.



polyurethane, rubber, paints and pigments, adhesives, and coatings.¹⁶⁹ Alkylphosphates such as tris(2-ethylhexyl) phosphate have found widespread usage in PVC and rubber by showing great low-temperature performances and low viscosity. Tributyl phosphate is also used for cellulose acetate and nitrocellulose. Low volatility and sustainable alternatives such as phosphonium-based ionic liquids have also shown good plasticizing properties.¹⁷⁰

2.3.7. Extraction ligand

2.3.7.1. Uranium extraction. P(v) compounds are essential in nuclear chemistry as extraction agents. Nuclear power plants typically rely on the fission of heavy actinide isotopes such as ²³⁵U to generate power. Uranium is generally obtained by purifying natural uranium ores according to different processes. According to the DAPEX (dialkyl phosphoric extraction) procedure, the uranium ore is treated with a strong acid such as nitric acid and then purified by solvent extraction, using a non-polar solvent such as kerosene and two organophosphorus extraction agents di-(2-ethylhexyl)phosphoric acid (Fig. 32(1)) with a 4% loading and tri-*n*-butyl phosphate (Fig. 32(2)) in an equivalent amount. Long alkyl substituents give those organophosphorus chemicals good lipophilicity and thus good solubility in inert solvents while coordinating selectively with uranyl ions, thanks to their high polarity. After complexation with di-(2-ethylhexyl)phosphoric acid (ligand L), uranyl complexes such as (UO₂)₂(NO₃)₂L₂ are formed.¹⁷¹ After separating the impurities by removing the aqueous phase, the uranium-rich kerosene phase is then treated with a strip solution containing either HCl, HF, or carbonate to transfer uranyl ions in the aqueous phase. The role of tributyl phosphate is thought to have a synergistic effect compared to using the phosphoric acid ligand alone: the tributyl phosphate additive is thought to contribute to the stability of (UO₂)₂(NO₃)₂L₂ complex by hydrogen bonding.¹⁷² Uranium dioxide will then be converted to UF₆ for isotopic enrichment before use as nuclear fuel.

Other organophosphorus compounds have been used for uranium extraction, such as trioctylphosphine oxide (Fig. 32(3)) and bis(2,4,4-trimethyl pentyl)phosphinic acid. A summary of all the extraction agents used for uranium in acidic solution and their comparison is given in a review written by Yoon and coworkers.¹⁷³ Phosphonate compounds such as dibutylbutyl phosphonate can be used as extractants as well.¹⁷⁴ As an alternative to uranium ores, uranium can also be extracted from phosphate rocks, which can be profitable when the demand is high. Since the 1990s, processes using both di-(2-ethylhexyl)phosphoric acid and trioctylphosphine oxides are considered state of the art concerning uranium extraction from phosphate rocks. A comprehensive review published in 2014 details the different processes and the challenges ahead for this unconventional source.¹⁷⁵

2.3.7.2. Recycling nuclear waste. After nuclear fuel usage, nuclear powerplants produce residual materials, which are mixtures of fission products and still valuable uranium and plutonium materials. Recovering those energy-carrier isotopes to recycle them for conversion, enrichment, and fuel usage is the central role of reprocessing plants. This strategy aims at a more economical and sustainable use of actinides in a context

where uranium resources are finite. The reprocessing plants use the PUREX (plutonium uranium extraction) process to produce UO₂(NO₃)₂ and Pu(NO₃)₄, which can either be re-enriched in the case of uranium or used in a mixed uranium–plutonium fuel called mixed oxide fuel (MOX) for plutonium. The PUREX process relies on liquid–liquid extraction to isolate useful actinides and separate them from wasteful fission products. After the dissolution of the actinides in nitric acid, the actinide nitrates are treated with kerosene and 30% tri-*n*-butyl phosphate, which binds selectively to uranium and plutonium of +IV and +VI oxidation state to form stable complexes.¹⁷⁶ The other fission by-products will remain in the aqueous layer. Plutonium complexes can later be separated from the uranium ones by using a reducing agent such as ferrous sulfamate or hydrazine to obtain plutonium species to a +III oxidation state, which would then migrate towards the aqueous phase.

2.3.7.3. Miscellaneous. In addition to uranium extraction, the use of P(v) compounds can be extended to the extraction of valuable rare-earth elements.^{177–182} Finally, more recent strategies include the design of pillarenes or calixarenes bearing phosphine oxide moieties to extract uranium from ionic liquids.^{183,184}

3. Synthesis of organophosphorus: classical pathways & associated challenges

After examining their different applications, we will explore in this section the industrial synthetic ways for fine organophosphorus chemicals and consider their associated limitations and sustainability issues. For clarity, the reactions shown here are the most used synthesis pathways towards non-nitrogen-containing organophosphorus compounds. Critical alternative pathways are also shown in dashed lines, but many other reactions have been overlooked when less relevant. For instance, the reductions of phosphine oxides to phosphines or phosphates to phosphites performed with lithium aluminium hydride have been disregarded as they show lesser industrial value. Therefore, the aim here is not to present an exhaustive list of all synthetic routes for organophosphorus molecules but rather to show the trends in a simplified manner, as described above (Fig. 33).

All the synthetic pathways to organophosphorus compounds start from the same universal intermediate, which is white phosphorus. By itself, molecular phosphorus does not have direct functionalization pathways used on an industrial scale towards useful organophosphorus. For example, both white phosphorus and PCl₃ do not have significant applications other than in the military field. However, white phosphorus can be converted to other intermediates: phosphorus trichloride, phosphine, and to a lesser extent, phosphorus pentoxide (anhydride form of pure orthophosphoric acid). The latter can be formed through the exothermic oxidation of white phosphorus and can subsequently provide monosubstituted and disubstituted acid phosphates after reaction with alcohols.

Phosphine PH₃ can be prepared by treating white phosphorus with water and either sodium or potassium hydroxide.





Fig. 33 Overview of synthetic routes towards organophosphorus compounds.

Phosphine is a dangerous and toxic gas by itself, it but can be converted to either primary or secondary phosphines through hydrophosphination of olefins (more details below). In the big picture, phosphorus trichloride remains the lowest common denominator as a ubiquitous intermediate for most organophosphorus. Its preparation necessitates the treatment of white phosphorus with toxic gaseous chlorine under reflux. Phosphorus trichloride can provide access to tertiary phosphines after alkylation through Grignard reagents. Subsequent oxidation by hydrogen peroxide or octasulfur S_8 yields tertiary phosphine oxides or sulfides. Treatment of PCl_3 with alcohol can yield either disubstituted or trisubstituted phosphites according to the presence or absence of base additives. Monoalkylation of PCl_3 occurs through a high-temperature reaction with alkanes, affording dichlorophosphines $PRCl_2$. Consecutive alcohol substitution gives phosphonites, whereas $PRCl_2$ disproportionation produces chlorophosphines PR_2Cl . Chlorophosphines also provide access to secondary phosphine oxides after oxidation or phosphinite after alcohol substitution. Finally, the oxidation of phosphorus trichloride can form another common intermediate: phosphoryl chloride $P(O)Cl_3$. Treatment with carbanion equivalent will afford tertiary phosphine oxides while substitution with alcohol will generate trisubstituted phosphates. Considering that P_4 , PCl_3 , and $POCl_3$ are ubiquitous P-intermediates, obvious sustainability issues stand out when looking at the current industrial procedures:

-Energy costly processes are necessary, such as the white phosphorus synthesis from apatite or the synthesis of PCl_3 .

-The use of toxic reagents chlorine gas is also indispensable, and phosphorus chlorides intermediates will provide hazardous halide by-products after functionalization.

-In terms of redox efficiency, the global synthetic route can also be perfected: the preparation of oxidized organophosphorus (oxidation number from +I to +V) such as phosphites, phosphonates, and phosphates require the reduction of phosphorus from naturally occurring apatite (+V) to molecular phosphorus (0) and then re-oxidation.

-In terms of redox economy,¹⁸⁵ P-intermediates with a more balanced oxidation number would be highly desirable in order to obtain both phosphines (-III) and phosphates (+V), rather than the currently used PCl_3 (+III) and $POCl_3$ (+V).

-Finally, all P-intermediates, namely P_4 , PCl_3 , and $POCl_3$, are eminently hazardous and uneasy to handle.

3.1. Towards more sustainable P-building blocks?

Considering the growing need for more sustainable chemical transformations, better alternatives complying with the 12 principles of green chemistry and concepts such as atom economy should be developed for the organophosphorus industry.^{186,187} Different strategies currently being investigated on a laboratory scale to solve this problem are highlighted in this section. The list of pathways presented below is far from exhaustive, and a more curious reader should refer to the few reviews already published on the subject.¹⁸⁸⁻¹⁹¹

3.1.1. White phosphorus activation. As molecular phosphorus is the first obligatory intermediate, one logical pathway to obtain organophosphorus would be through direct P_4 -activation. However, nucleophiles such as aqueous hydroxide ions are known to disrupt the tetrahedral architecture of P_4 , thus furnishing either reactive phosphine or phosphide ions which can react in the presence of electrophiles *in situ*.¹⁹² Treatment of molten white phosphorus with potassium hydroxide within



a biphasic system in the presence of alkyl halides and a phase-transfer catalyst affords alkylphosphonic acids.^{193,194} Primary phosphines can be synthesized from alkyl halides by treating the same molten white phosphorus with organolithium reagents and following an aqueous work-up.¹⁹⁵ By playing with conditions or using Grignard reagents, organosodiums, organozincs, or various polysubstituted phosphines can be produced in the presence of alkyl halides.^{196,197} An interesting procedure showed that phenol derivatives react with P_{white} under aerobic conditions, with the help of iron complexes and iodine to form arylphosphates.¹⁹⁸ White phosphorus can also be functionalized through a radical mechanism by acting as a radical trap: photogenerated carbon radicals can react with P_{white} to afford alkylphosphonic acid after hydrolysis and oxidation by hydrogen peroxide.¹⁹⁹ A more recent radical reaction reported is the synthesis of tertiary phosphines from alkyl halides and P_{white} using titanium complexes as catalysts.²⁰⁰ Transition metal catalysis can also be used: organorhodiums generate primary phosphines from white phosphorus in low yields after hydrogenation.²⁰¹ Electrochemical protocols showed that P_{white} could be converted to phosphine PH_3 in moderate to good yields, trialkylphosphates, phosphoramides, or even triaryl phosphines.^{202–204} Most protocols are summarized above (Fig. 34). The essential drawback of these methods is that not all phosphorus atoms can be activated, causing a poor atom economy. Indeed, as the tetrahedral structure is broken down, the reactivity of the P–P bonds diminishes. In addition, white phosphorus is both toxic and remains pyrophoric under air at temperatures above 34 °C, making it a very hazardous source of phosphorus. Furthermore, it bears a very poor solubility, except when using carbon disulfide as a solvent.

3.1.2. Red phosphorus activation. Red phosphorus, obtained by either heating white phosphorus at 300 °C in the absence of oxygen or by exposing it to sunlight, is a markedly safer alternative as it does not spontaneously ignite under air at

temperatures below 240 °C. Despite its usage as an inorganic flame retardant and its recently discovered photocatalytic activities,^{205–207} P_{red} can also be activated to synthesize organophosphorus compounds, although its polymeric structure makes it less reactive than P_{white} . P_{red} can give trialkylphosphine oxides after basic treatment in biphasic systems in the presence of alkyl halides.^{208,209} Another method using nickel bromide as a transition metal catalyst and aryl iodide provides triarylphosphine oxides after oxidation.²¹⁰ Primary phosphines can be obtained in moderate yields through the reaction of P_{red} with haloalkane in the presence of sodium metal and *t*-butyl alcohol.²¹¹ A similar procedure gives tertiary phosphines from alkenes.²¹² Hydrophosphonylation of aldehydes can be performed by red phosphorus in a basic aqueous medium to produce hydroxyphosphonites.²¹³ The same article reports the synthesis of alkylphosphonic acids by treating P_{red} with aqueous hydroiodic acid in dioxane in the presence of an aldehyde. Several procedures are detailed below (Fig. 35). Finally, many procedures which are not detailed here use superbasic conditions (KOH/DMSO) for the hydrophosphorylation of olefins and alkynes, providing pathways to phosphine oxides, phosphinic acids, and phosphines from either P_{red} , nanocomposites of P_{red} or P_{white} .²¹⁴ Nonetheless, akin to P_{white} , problems linked to difficult solubility and a wasteful stoichiometry perdure.

3.1.3. Phosphine activation. As previously stated, PH_3 is mostly well known for its reactions with alkenes to form secondary and secondary phosphines through hydrophosphination, primarily for linear alkenes.²¹⁵ This reaction can be acid or base-catalyzed in the case of Michael acceptors or even under neat conditions in the absence of solvents.²¹⁶ Hydrophosphination of imines and carbonyls have also been reported, and addition onto alkynes give vinylphosphines. Free-radical processes initiated by either radical initiators or UV irradiation can provide an efficient pathway for synthesizing tertiary phosphines, as seen in the industrial preparation of tributylphosphine and trioctylphosphines.²¹⁷ In the presence of strong bases (KOH) or alkali/alkaline earth metals (Li, Na, K, or



Fig. 34 White phosphorus as an intermediate for organophosphorus synthesis.

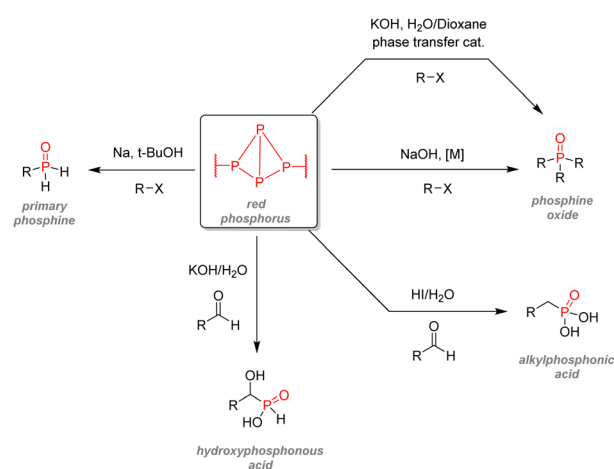


Fig. 35 Red phosphorus as an intermediate for organophosphorus synthesis.



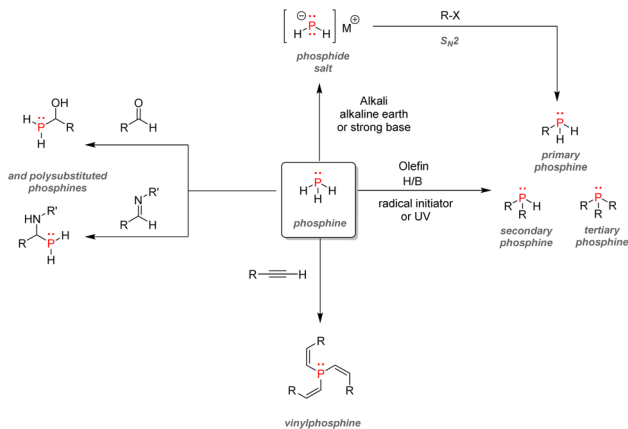


Fig. 36 Phosphine as an intermediate for organophosphorus synthesis.

Ca), nucleophilic phosphide salts can be prepared from PH_3 . Subsequent alkylation in the presence of alkyl halides can also provide a path towards primary and secondary phosphines, albeit less atom economical (Fig. 36). Because PH_3 is a toxic and flammable gas, exuding highly unpleasant odour, phosphine is not an easy reagent to store or handle, thus making it a poor candidate for a more sustainable intermediate for the synthesis of organophosphorus compounds.

3.1.4. Phosphoric acid. Instead of relying on P_4 as a central intermediate, which carries inherent drawbacks such as requiring energy-intensive processes and handling hazardous and pyrophoric white phosphorus material, one radical alternative is to bypass P_4 altogether and directly functionalize existing inorganic sources of $\text{P}(\text{v})$. This paradigm shift would

avoid numerous sustainability issues linked with the traditional use of P_4 , PCl_3 and POCl_3 , and could potentially simplify the organophosphorus industry. As most phosphate rocks are converted to orthophosphoric acid through the ‘wet process’ as previously stated, using this H_3PO_4 and its already available technology can be a good starting point. In addition, orthophosphoric acid can be used as a $\text{P}(\text{v})$ source to form a mono-ester of phosphoric acids (or monosubstituted phosphates) by esterification. The procedures were developed in the 1950s, notably for the phosphorylation of nucleoside phosphates. The use of phosphorus pentoxide, polyphosphoric acids, or their conjugate bases leads to synthesizing a mixture of mono- and diesters.^{218–224} But to this date, direct functionalization from orthophosphoric acid is still limited.

3.1.5. Bis(trichlorosilyl)phosphide salts. Another opportunity relies on the conversion of orthophosphoric acid to more reactive intermediates. Inspired by the strong reducing properties of trichlorosilane – an already established and commonly synthesized precursor to ultrapure silicon for semiconductors – the Cummins group showed an extremely promising bis(trichlorosilyl)phosphide salt intermediate that could be formed when using trichlorosilane to reduce trimetaphosphate salts.²²⁵ Those salts can be obtained after simple dehydration of orthophosphoric acid using sodium chloride before an anion exchange step.^{226,227} The afforded bis(trichlorosilyl)phosphide salt can then be readily converted either to mono- or disubstituted phosphines or even phosphonium salts when reacting with alkyl halides, to PH_3 after treatment with water or to hexafluorophosphate anion when using XeF_2 (Fig. 37). A more recent report by the same group showed that the synthesis of monosubstituted phosphine could be carried out as a one-pot process starting directly from orthophosphoric acid.²²⁸ A

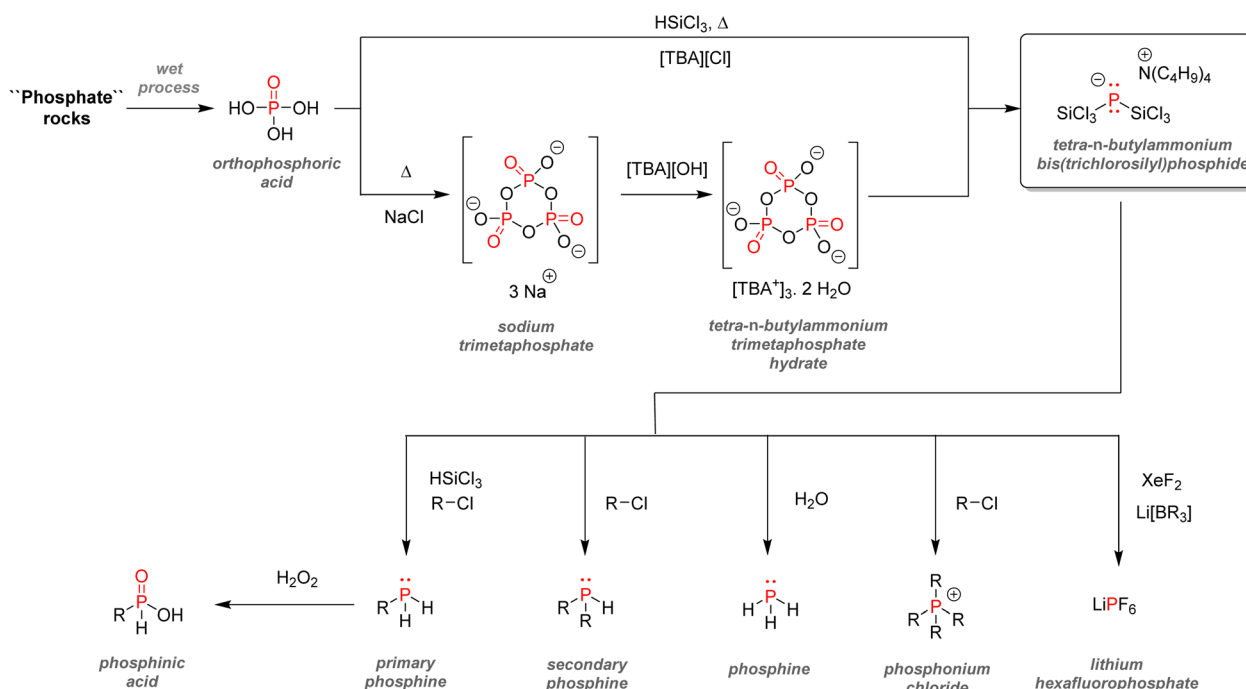


Fig. 37 Bis(trichlorosilyl)phosphide salts as intermediates for organophosphorus synthesis.





Fig. 38 Hypophosphites as intermediates for organophosphorus synthesis.

multi-step synthesis of Fosinopril (a phosphinate therapeutic used against heart conditions) was also undertaken, using bis(trichlorosilyl)phosphide as the phosphorus source. This strategy is a promising substitute to phosphine as a phosphorus platform intermediate: the handling of white phosphorus is avoided, and many synthetic routes towards low oxidation state phosphorus species, such as trivalent P(III) ones, have been developed.

3.1.6. Hypophosphite. Another extremely encouraging lead is the pioneering work established by Montchamp. His laboratory aims at substituting PCl_3 and PH_3 with nontoxic and easier to handle hypophosphite derivatives. The advantages of this strategy are that a hypophosphite supply is directly available industrially, as sodium phosphinate salts can be obtained through alkaline hydrolysis of white phosphorus. Besides, hypophosphites bear an oxidation number (+I), which is right in the middle compared to phosphates (+V) and phosphines (−III).²²⁹ Moreover, sodium phosphinate is solid at room temperature and water-soluble. Alkyl hypophosphites called hypophosphites can be prepared *via* either alkoxy silane-mediated esterification or Fischer-type esterification of phosphinate salts.^{230,231} A wide range of reactions has been developed to alkylate hypophosphites to H-phosphinates, including nucleophilic reactions, radical reactions, hydrophosphinylations, benzylations, allylations, or cross-coupling reactions. In turn, those H-phosphinates can be converted to phosphonates, thiophosphonates, phosphinates, *etc.*^{187,232–234} Hypophosphorous acid H_3PO_2 (the conjugate acid of phosphinate salt) can also generate useful disubstituted phosphites using nickel chloride or nickel on silica as catalysts and various alcohols, thus completely bypassing the use of PCl_3 .²³⁵ The different pathways are summarized below (Fig. 38). Although this hypophosphite pathway needs a preliminary conversion of white phosphorus, numerous useful synthetic routes have been developed towards high oxidation state phosphorus and P(V) species.

4. Conclusion

With such a rich diversity in their chemical properties and their wide range of applications in many significant processes,

organophosphorus compounds are undoubtedly a valuable class of chemicals worthy of further research and development. Indeed, agriculture, flame retardants, therapeutics, or optoelectronic materials are all fields related in one way or another to the phosphorus industry. However, the latter presently relies on the extraction and transformation of fossil resources unequally distributed around the globe. Furthermore, the future scarcity of phosphate rocks during a potential peak phosphorus event will mechanically damage human activities and raise international tensions by making phosphorus a strategic commodity. As such, efforts to develop a circular economy for phosphorus should be promoted to make this resource less precarious, more sustainable and reduce excess phosphate runoff towards the ocean, which could endanger marine ecosystems and fishing activities. Another issue is linked to the specific chemical processes employed to functionalize organophosphorus: most synthetic routes depend on hazardous P–Cl containing building blocks. They also rely heavily on Grignard or organolithium alkylating agents and energetically costly transformations while creating potentially hazardous halogenated by-products. With the ongoing environmental concern and drive to develop more sustainable processes, a more holistic approach toward the organophosphorus industry is needed: develop safer building blocks and milder alternative procedures featuring a better atom economy and redox efficiency. The phosphide salt route, the hypophosphite route, and the future developments of more efficient P(V) to P(III) reduction or the direct functionalization of phosphoric acid are among the promising new pathways that navigate organophosphorus chemistry with a better redox economy and sustainability.

Author contributions

Sosthène Ung wrote the review, Prof. C.-J. Li proofread the manuscript.

Conflicts of interest

There are no conflicts to declare.



Acknowledgements

The authors acknowledge Prof. Damha, Prof. Tsantrizos, Prof. Moitessier from the department of Chemistry of McGill University and Prof. Virieux from École Nationale Supérieure de Chimie de Montpellier for their advice and guidance.

References

- J. D. Watson and F. H. C. Crick, *Nature*, 1953, **171**, 737–738.
- E. Gajewski, D. K. Steckler and R. N. Goldberg, *J. Biol. Chem.*, 1986, **261**, 12733–12737.
- P. Cohen, *Nat. Cell Biol.*, 2002, **4**, E127–E130.
- M. E. Weeks, *J. Chem. Educ.*, 1933, **10**, 302.
- P. Gengembre, *Mém. Mathématique Phys.*, 1783, **10**, 651–658.
- J. L. Gay-Lussac and L. J. Thénard, *Gaz. Natl. Ou Monit. Univers.*, 1808, **40**, 581–582.
- M. F. Crass, *J. Chem. Educ.*, 1941, **18**, 277.
- J. P. W. Hughes, R. Baron, D. H. Buckland, M. A. Cooke, J. D. Craig, D. P. Duffield, A. W. Grosart, P. W. J. Parkes and A. Porter, *Occup. Environ. Med.*, 1962, **19**, 83–99.
- D. M. Ivell, *Procedia Eng.*, 2012, **46**, 166–171.
- J. B. Readman, *US Pat.*, US417943A, p. 1889.
- H. Rose, *Ann. Chim. Phys.*, 1832, **51**, 1–62.
- A. W. Hofmann, *Justus Liebigs Ann. Chem.*, 1857, **103**, 357–358.
- A. Cahours and H. Gal, *Justus Liebigs Ann. Chem.*, 1870, **155**, 223–230.
- A. Cahours and H. Gal, *Justus Liebigs Ann. Chem.*, 1870, **155**, 355–359.
- G. A. Petroianu, *Pharmazie*, 2010, **4**, 306–311.
- B. A. Arbusow, *Pure Appl. Chem.*, 1964, **9**, 307–336.
- A. Michaelis and R. Kaehne, *Ber. Dtsch. Chem. Ges.*, 1898, **31**, 1048–1055.
- D. E. C. Corbridge, in *Phosphorus: Chemistry, Biochemistry and Technology*, Taylor & Francis, Boca Raton, 6th edn, 2013, p. 5.
- F. Matsumura, in *Toxicology of Insecticides*, ed. F. Matsumura, Springer US, Boston, MA, 1985, pp. 45–109.
- J. B. Tucker, *War of Nerves: Chemical Warfare from World War I to Al-Qaeda*, Anchor Books, First Anchor Books ed., New York, NY, 2007.
- Chemical Weapons Convention of 1993*, <https://sites.fas.harvard.edu/~hsp/cwc/cwcbyart.html>, accessed 15 July 2020.
- D. Darling, *Terrestrial abundance of elements*, <https://www.daviddarling.info/encyclopedia/E/eltterr.html>, accessed 12 July 2020.
- J. R. Rumble, D. R. Lide and T. J. Bruno, *CRC Handbook of Chemistry and Physics: a Ready-Reference Book of Chemical and Physical Data*, CRC Press, Boca Raton, 98th edn, 2017.
- K. Barbalace, *Periodic Table of Elements: Phosphorus – P*, <https://environmentalchemistry.com/yogi/periodic/P.html>, accessed 12 July 2020.
- I. Hanukoglu, *List of Elements of the Periodic Table – Sorted by Abundance in Earth's crust*, <https://www.science.co.il/elements/?s=Earth>, accessed 12 July 2020.
- Z. Yuan, S. Jiang, H. Sheng, X. Liu, H. Hua, X. Liu and Y. Zhang, *Environ. Sci. Technol.*, 2018, **52**, 2438–2450.
- M. Chen and T. E. Graedel, *Glob. Environ. Change*, 2016, **36**, 139–152.
- D. Cordell, J.-O. Drangert and S. White, *Glob. Environ. Change*, 2009, **19**, 292–305.
- D. Bernhardt and J. F. Reilly II, *Mineral Commodity Summaries 2020*, U.S. Geological Survey, Reston, VA, 2020.
- K. Schrödter, G. Bettermann, T. Staffel, T. Klein and T. Hofmann, Phosphoric Acid and Phosphates, in *Ullmann's Encyclopedia of Industrial Chemistry*, American Chemical Society, 2000.
- G. Villalba, Y. Liu, H. Schroder and R. U. Ayres, *J. Ind. Ecol.*, 2008, **12**, 557–569.
- Y. I. Asa, *Fertilizer Industry Handbook 2018*, 2018, p. 97.
- H. Diskowski and T. Hofmann, Phosphorus. in *Ullmann's Encyclopedia of Industrial Chemistry*, American Chemical Society, 2000.
- D. E. C. Corbridge, in *Phosphorus: Chemistry, Biochemistry and Technology*, Taylor & Francis, Boca Raton, 6th edn, 2013, pp. 93–94.
- G. Bettermann, K. Werner, G. Riess and T. Hofmann, Phosphorus Compounds, Inorganic, in *Ullmann's Encyclopedia of Industrial Chemistry*, American Chemical Society, 2000.
- N. N. Greenwood and A. Earnshaw, in *Chemistry of the Elements*, Butterworth-Heinemann, Oxford; Boston, 2nd edn, 1997, p. 495.
- E. Wiberg, N. Wiberg and A. F. Holleman, *Inorganic Chemistry*, Academic Press; De Gruyter, San Diego; Berlin; New York, 1st edn, 2001.
- J. Svara, N. Weferling and T. Hofmann, Phosphorus Compounds, Organic, in *Ullmann's Encyclopedia of Industrial Chemistry*, American Chemical Society, 2006.
- L. M. E. Percival, D. P. G. Bond, M. Rakociński, L. Marynowski, A. v. S. Hood, T. Adatte, J. E. Spangenberg and K. B. Föllmi, *Glob. Planet. Change*, 2020, **184**, 103070.
- D. J. Conley, S. Björck, E. Bonsdorff, J. Carstensen, G. Destouni, B. G. Gustafsson, S. Hietanen, M. Kortekaas, H. Kuosa, H. E. Markus Meier, B. Müller-Karulis, K. Nordberg, A. Norkko, G. Nürnberg, H. Pitkänen, N. N. Rabalais, R. Rosenberg, O. P. Savchuk, C. P. Slomp, M. Voss, F. Wulff and L. Zillén, *Environ. Sci. Technol.*, 2009, **43**, 3412–3420.
- A. J. Watson, T. M. Lenton and B. J. W. Mills, *Philos. Trans. R. Soc., A*, 2017, **375**, 2102.
- A. M. Michalak, E. J. Anderson, D. Beletsky, S. Boland, N. S. Bosch, T. B. Bridgeman, J. D. Chaffin, K. Cho, R. Confesor, I. Daloğlu, J. V. DePinto, M. A. Evans, G. L. Fahnenstiel, L. He, J. C. Ho, L. Jenkins, T. H. Johengen, K. C. Kuo, E. LaPorte, X. Liu, M. R. McWilliams, M. R. Moore, D. J. Posselt, R. P. Richards, D. Scavia, A. L. Steiner, E. Verhamme, D. M. Wright and M. A. Zagorski, *Proc. Natl. Acad. Sci. U. S. A.*, 2013, **110**, 6448–6452.
- J. Rockström, W. Steffen, K. Noone, Å. Persson, F. S. Chapin, E. F. Lambin, T. M. Lenton, M. Scheffer,



- C. Folke, H. J. Schellnhuber, B. Nykvist, C. A. de Wit, T. Hughes, S. van der Leeuw, H. Rodhe, S. Sörlin, P. K. Snyder, R. Costanza, U. Svedin, M. Falkenmark, L. Karlberg, R. W. Corell, V. J. Fabry, J. Hansen, B. Walker, D. Liverman, K. Richardson, P. Crutzen and J. A. Foley, *Nature*, 2009, **461**, 472–475.
- 44 A. R. Jupp, S. Beijer, G. C. Narain, W. Schipper and J. C. Slootweg, *Chem. Soc. Rev.*, 2021, **50**, 87–101.
- 45 S. Marque and P. Tordo, *Reactivity of Phosphorus Centered Radicals*, Springer, Berlin Heidelberg, 2005.
- 46 A. C. Gaumont and J. M. Denis, *Chem. Rev.*, 1994, **94**, 1413–1439.
- 47 E. H. Krenske, K. N. Houk, B. A. Arndtsen and D. J. S. Cyr, *J. Am. Chem. Soc.*, 2008, **130**, 10052–10053.
- 48 N. L. Dunn, M. Ha and A. T. Radosevich, *J. Am. Chem. Soc.*, 2012, **134**, 11330–11333.
- 49 T. E. Gier, *J. Am. Chem. Soc.*, 1961, **83**, 1769–1770.
- 50 F. Mathey, in *Modern Heterocyclic Chemistry*, John Wiley & Sons, Ltd, 2011, pp. 2071–2116.
- 51 R. D. Baechler and K. Mislow, *J. Am. Chem. Soc.*, 1970, **92**, 3090–3093.
- 52 R. H. Crabtree, *The Organometallic Chemistry of the Transition Metals*, Wiley, Hoboken, New Jersey, 6th edn, 2014.
- 53 C. A. Tolman, *Chem. Rev.*, 1977, **77**, 313–348.
- 54 C. A. Tolman, *J. Am. Chem. Soc.*, 1970, **92**, 2956–2965.
- 55 P. Dierkes and P. W. N. M. van Leeuwen, *J. Chem. Soc., Dalton Trans.*, 1999, 1519–1530.
- 56 N. Miyaura, T. Ishiyama, H. Sasaki, M. Ishikawa, M. Sato and A. Suzuki, *J. Am. Chem. Soc.*, 1989, **111**, 314–321.
- 57 S. R. Chemler, D. Trauner and S. J. Danishefsky, *Angew. Chem., Int. Ed.*, 2001, **40**, 4544–4568.
- 58 M. R. Netherton, C. Dai, K. Neuschütz and G. C. Fu, *J. Am. Chem. Soc.*, 2001, **123**, 10099–10100.
- 59 R. Noyori and H. Takaya, *Acc. Chem. Res.*, 1990, **23**, 345–350.
- 60 T. Ohta, H. Takaya and R. Noyori, *Inorg. Chem.*, 1988, **27**, 566–569.
- 61 R. Noyori, M. Tokunaga and M. Kitamura, *Bull. Chem. Soc. Jpn.*, 1995, **68**, 36–55.
- 62 W. S. Knowles, *Acc. Chem. Res.*, 1983, **16**, 106–112.
- 63 S. Akutagawa, *Appl. Catal., A*, 1995, **128**, 171–207.
- 64 D. Morales-Morales and C. M. Jensen, *The Chemistry of Pincer Compounds*, Elsevier, Amsterdam, Netherlands, 2007.
- 65 D. Morales-Morales, *Pincer Compounds: Chemistry and Applications*, Elsevier, Amsterdam, Netherlands, 2018.
- 66 R. A. Gossage and G. van Koten, *The Privileged Pincer-Metal Platform: Coordination Chemistry & Applications*, Springer International Publishing, Cham, 1st edn, 2016.
- 67 G. van Koten and D. Milstein, *Organometallic Pincer Chemistry*, Springer Berlin, Heidelberg, 2013.
- 68 K. J. Szabó and O. F. Wendt, *Pincer and Pincer-type Complexes: Applications in Organic Synthesis and Catalysis*, Wiley-VCH, Weinheim, 2014.
- 69 C. Gunanathan, Y. Ben-David and D. Milstein, *Science*, 2007, **317**, 790–792.
- 70 H. Valdés, M. A. García-Eleno, D. Canseco-Gonzalez and D. Morales-Morales, *ChemCatChem*, 2018, **10**, 3136–3172.
- 71 M. Feller, C. Gunanathan, A. Kumar, R. Langer, M. Montag, T. Schaub, M. Vogt and T. Zell, David Milstein: Shaping Organometallic Catalysis Over Five Decades, *ChemViews*, 2022.
- 72 C. Li, J. Kan, Z. Qiu, J. Li, L. Lv and C. Li, *Angew. Chem., Int. Ed.*, 2020, **59**, 4544–4549.
- 73 M. Asay and D. Morales-Morales, *Dalton Trans.*, 2015, **44**, 17432–17447.
- 74 H. Valdés, J. M. Germán-Acacio, G. van Koten and D. Morales-Morales, *Dalton Trans.*, 2022, **51**, 1724–1744.
- 75 D. W. Stephan and G. Erker, *Angew. Chem., Int. Ed.*, 2015, **54**, 6400–6441.
- 76 A. R. Jupp and D. W. Stephan, *Trends Chem.*, 2019, **1**, 35–48.
- 77 V. V. Grushin, *Chem. Rev.*, 2004, **104**, 1629–1662.
- 78 T. M. Shaikh, C.-M. Weng and F.-E. Hong, *Coord. Chem. Rev.*, 2012, **256**, 771–803.
- 79 L. D. Quin, *A Guide to Organophosphorus Chemistry*, Wiley, New York, 2000.
- 80 D. B. Chesnut, *J. Phys. Chem. A*, 2003, **107**, 4307–4313.
- 81 D. B. Chesnut and A. Savin, *J. Am. Chem. Soc.*, 1999, **121**, 2335–2336.
- 82 A. E. Reed and P. v. R. Schleyer, *J. Am. Chem. Soc.*, 1990, **112**, 1434–1445.
- 83 E. N. Tsvetkov and A. A. Korkin, *Theor. Exp. Chem.*, 1985, **21**, 36–42.
- 84 K. M. Diemoz and A. K. Franz, *J. Org. Chem.*, 2019, **84**, 1126–1138.
- 85 B. Klabuhn, H. Goetz, P. Steirl and D. Alschner, *Tetrahedron*, 1976, **32**, 603–607.
- 86 A. W. Johnson and H. L. Jones, *J. Am. Chem. Soc.*, 1968, **90**, 5232–5236.
- 87 K. Gluza and P. Kafarski, in *Drug Discovery*, ed. H. El-Shemy, InTech, 2013.
- 88 P. Finkbeiner, J. P. Hehn and C. Gnam, *J. Med. Chem.*, 2020, **63**, 7081–7107.
- 89 A. Emadi, R. J. Jones and R. A. Brodsky, *Nat. Rev. Clin. Oncol.*, 2009, **6**, 638–647.
- 90 W. H. Organization, *Model List of Essential Medicines: 21st List*, 2019, <https://www.who.int/publications/i/item/WHOMVPEMPIAU2019.06>, accessed 2020-07-31.
- 91 A. C. Dijkmans, N. V. O. Zacarias, J. Burggraaf, J. W. Mouton, E. Wilms, C. van Nieuwkoop, D. J. Touw, J. Stevens and I. M. C. Kamerling, *Antibiotics*, 2017, **6**, 24.
- 92 E. Helgstrand, B. Eriksson, N. G. Johansson, B. Lannero, A. Larsson, A. Misiorny, J. O. Noren, B. Sjöberg, K. Stenberg, G. Stening, S. Stridh and B. Oberg, *Science*, 1978, **201**, 819–821.
- 93 A. Canestri, J. Ghosn, M. Wirten, F. Marguet, N. Ktorza, I. Boubezari, S. Dominguez, P. Bossi, E. Caumes, V. Calvez and C. Katlama, *Antiviral Ther.*, 2006, **11**, 561–566.
- 94 E. M. Lewiecki, *Ther. Adv. Chronic Dis.*, 2010, **1**, 115–128.
- 95 M. T. Drake, B. L. Clarke and S. Khosla, *Mayo Clin. Proc.*, 2008, **83**, 1032–1045.
- 96 K. S. E. Tanaka, E. Dietrich, S. Ciblat, C. Métayer, F. F. Arhin, I. Sarmiento, G. Moeck, T. R. Parr and A. R. Far, *Bioorg. Med. Chem. Lett.*, 2010, **20**, 1355–1359.



- 97 B. Badet, K. Inagaki, K. Soda and C. T. Walsh, *Time-dependent inhibition of Bacillus stearothermophilus alanine racemase by (1-aminoethyl)phosphonate isomers by isomerization to noncovalent slowly dissociating enzyme-(1-aminoethyl)phosphonate complexes*, <https://pubs.acs.org/doi/pdf/10.1021/bi00359a029>, accessed 31 July 2020.
- 98 J. T. Whitteck, W. Ni, B. M. Griffin, A. C. Eliot, P. M. Thomas, N. L. Kelleher, W. W. Metcalf and W. A. van der Donk, *Angew. Chem., Int. Ed.*, 2007, **46**, 9089–9092.
- 99 J. Krapcho, C. Turk, D. W. Cushman, J. R. Powell, J. M. DeForrest, E. R. Spitzmiller, D. S. Karanewsky, M. Duggan and G. Rovnyak, *J. Med. Chem.*, 1988, **31**, 1148–1160.
- 100 E. Skordalakes, G. G. Dodson, D. S. C. Green, C. A. Goodwin, M. F. Scully, H. R. Hudson, V. V. Kakkar and J. J. Deadman, *J. Mol. Biol.*, 2001, **311**, 549–555.
- 101 R. Bone, N. S. Sampson, P. A. Bartlett and D. A. Agard, *Biochemistry*, 1991, **30**, 2263–2272.
- 102 A. Mucha, P. Kafarski and Ł. Berlicki, *J. Med. Chem.*, 2011, **54**, 5955–5980.
- 103 T. Usui, Y. Oiso, A. Tomita, E. Ogata, T. Uchida, K. Ikeda, T. Watanabe and S. Higuchi, *Int. J. Clin. Pharmacol. Ther.*, 1997, **35**, 239–244.
- 104 C. M. Shipman, P. I. Croucher, R. G. Russell, M. H. Helfrich and M. J. Rogers, *Cancer Res.*, 1998, **58**, 5294–5297.
- 105 P. Anderson, *Expert Opin. Pharmacother.*, 2006, **7**, 1475–1486.
- 106 M. Fellner, R. P. Baum, V. Kubíček, P. Hermann, I. Lukeš, V. Prasad and F. Rösch, *Eur. J. Nucl. Med. Mol. Imaging*, 2010, **37**, 834.
- 107 E. D. Clercq, *Nat. Rev. Microbiol.*, 2004, **2**, 704–720.
- 108 D. Siegel, H. C. Hui, E. Doerffler, M. O. Clarke, K. Chun, L. Zhang, S. Neville, E. Carra, W. Lew, B. Ross, Q. Wang, L. Wolfe, R. Jordan, V. Soloveva, J. Knox, J. Perry, M. Perron, K. M. Stray, O. Barauskas, J. Y. Feng, Y. Xu, G. Lee, A. L. Rheingold, A. S. Ray, R. Bannister, R. Strickley, S. Swaminathan, W. A. Lee, S. Bavari, T. Cihlar, M. K. Lo, T. K. Warren and R. L. Mackman, *J. Med. Chem.*, 2017, **60**, 1648–1661.
- 109 R. T. Eastman, J. S. Roth, K. R. Brimacombe, A. Simeonov, M. Shen, S. Patnaik and M. D. Hall, *ACS Cent. Sci.*, 2020, **6**, 672–683.
- 110 A. Fung, Z. Jin, N. Dyatkina, G. Wang, L. Beigelman and J. Deval, *Antimicrob. Agents Chemother.*, 2014, **58**, 3636–3645.
- 111 S. Safrin, J. Cherrington and H. S. Jaffe, *Rev. Med. Virol.*, 1997, **7**, 145–156.
- 112 C. P. Holstege, M. Kirk and F. R. Sidell, *Crit. Care Clin.*, 1997, **13**, 923–942.
- 113 F. Matsumura, *Toxicology of Insecticides*, Springer US, Boston, MA, 1985.
- 114 R. L. Metcalf (deceased) and A. R. Horowitz, Insect Control, 2. Individual Insecticides, in *Ullmann's Encyclopedia of Industrial Chemistry*, American Cancer Society, 2014, pp. 1–94.
- 115 D. Atwood and C. Paisley-Jones, *Pesticides Industry Sales and Usage 2008-2012 Market Estimates*, U.S. environmental Protection Agency, 2017, <https://www.epa.gov/pesticides/pesticides-industry-sales-and-usage-2008-2012-market-estimates>, accessed 2020-07-31.
- 116 R. L. Metcalf, Insect Control, in *Ullmann's Encyclopedia of Industrial Chemistry*, American Cancer Society, 2000.
- 117 M. Eddleston, N. A. Buckley, P. Eyer and A. H. Dawson, *Lancet*, 2008, **371**, 597–607.
- 118 S. Berg and E. A. Bittner, *The MGH Review of Critical Care Medicine*, Lippincott Williams & Wilkins, 2013.
- 119 M. A. Dar, G. Kaushik and J. F. Villareal Chiu, in *Abatement of Environmental Pollutants*, ed. P. Singh, A. Kumar and A. Borthakur, Elsevier, 2020, pp. 25–66.
- 120 F. Müller and A. P. Applebyki, Weed Control, 2. Individual Herbicides. in *Ullmann's Encyclopedia of Industrial Chemistry*, American Cancer Society, 2011.
- 121 S. O. Duke and S. B. Powles, *Pest Manag. Sci.*, 2008, **64**, 319–325.
- 122 C. Tomlin, *The Pesticide Manual: a World Compendium*, British Crop Protection Council, Alton, Hampshire, 15th edn, 2009.
- 123 B. K. Singh and A. Walker, *FEMS Microbiol. Rev.*, 2006, **30**, 428–471.
- 124 S. M. Bradberry, A. T. Proudfoot and J. A. Vale, *Toxicol. Rev.*, 2004, **23**, 159–167.
- 125 D. Cressey, Widely used herbicide linked to cancer, *Nat. News*, 2015, DOI: [10.1038/nature.2015.17181](https://doi.org/10.1038/nature.2015.17181).
- 126 L. N. Vandenberg, B. Blumberg, M. N. Antoniou, C. M. Benbrook, L. Carroll, T. Colborn, L. G. Everett, M. Hansen, P. J. Landrigan, B. P. Lanphear, R. Mesnage, F. S. vom Saal, W. V. Welshons and J. P. Myers, *J. Epidemiol. Community Health*, 2017, **71**, 613–618.
- 127 S. O. Duke, *Weed Sci.*, 2020, **68**, 201–207.
- 128 D. Schwartz, S. Berger, E. Heinzelmann, K. Muschko, K. Welzel and W. Wohlleben, *Appl. Environ. Microbiol.*, 2004, **70**, 7093–7102.
- 129 Y. Tao, C. Yang and J. Qin, *Chem. Soc. Rev.*, 2011, **40**, 2943–2970.
- 130 S. O. Jeon, K. S. Yook, C. W. Joo and J. Y. Lee, *Appl. Phys. Lett.*, 2009, **94**, 013301.
- 131 S. O. Jeon, K. S. Yook, C. W. Joo and J. Y. Lee, *Adv. Mater.*, 2010, **22**, 1872–1876.
- 132 S. Shao, J. Ding, T. Ye, Z. Xie, L. Wang, X. Jing and F. Wang, *Adv. Mater.*, 2011, **23**, 3570–3574.
- 133 C. Liu, Y. Li, Y. Li, C. Yang, H. Wu, J. Qin and Y. Cao, *Chem. Mater.*, 2013, **25**, 3320–3327.
- 134 C. Liu, Y. Gu, Q. Fu, N. Sun, C. Zhong, D. Ma, J. Qin and C. Yang, *Chem.–Eur. J.*, 2012, **18**, 13828–13835.
- 135 S. Shao, J. Ding, L. Wang, X. Jing and F. Wang, *J. Am. Chem. Soc.*, 2012, **134**, 20290–20293.
- 136 S. O. Jeon and J. Y. Lee, *J. Mater. Chem.*, 2012, **22**, 4233–4243.
- 137 N. Lin, J. Qiao, L. Duan, H. Li, L. Wang and Y. Qiu, *J. Phys. Chem. C*, 2012, **116**, 19451–19457.
- 138 N. Lin, J. Qiao, L. Duan, L. Wang and Y. Qiu, *J. Phys. Chem. C*, 2014, **118**, 7569–7578.



- 139 M. P. Duffy, W. Delaunay, P.-A. Bouit and M. Hissler, *Chem. Soc. Rev.*, 2016, **45**, 5296–5310.
- 140 C.-H. Lin, C.-W. Hsu, J.-L. Liao, Y.-M. Cheng, Y. Chi, T.-Y. Lin, M.-W. Chung, P.-T. Chou, G.-H. Lee, C.-H. Chang, C.-Y. Shih and C.-L. Ho, *J. Mater. Chem.*, 2012, **22**, 10684–10694.
- 141 D. Joly, D. Tondelier, V. Deborde, B. Geffroy, M. Hissler and R. Réau, *New J. Chem.*, 2010, **34**, 1603–1611.
- 142 Y. Matano, Y. Hayashi, H. Nakano and H. Imahori, *Heteroat. Chem.*, 2014, **25**, 533–547.
- 143 A. Kira, Y. Shibano, S. Kang, H. Hayashi, T. Umeyama, Y. Matano and H. Imahori, *Chem. Lett.*, 2010, **39**, 448–450.
- 144 C. Reus, M. Stolar, J. Vanderkley, J. Nebauer and T. Baumgartner, *J. Am. Chem. Soc.*, 2015, **137**, 11710–11717.
- 145 M. Stolar, J. Borau-Garcia, M. Toonen and T. Baumgartner, *J. Am. Chem. Soc.*, 2015, **137**, 3366–3371.
- 146 C. Reus and T. Baumgartner, *Dalton Trans.*, 2016, **45**, 1850–1855.
- 147 X.-D. Jiang, J. Zhao, D. Xi, H. Yu, J. Guan, S. Li, C.-L. Sun and L.-J. Xiao, *Chem.–Eur. J.*, 2015, **21**, 6079–6082.
- 148 X. Chai, X. Cui, B. Wang, F. Yang, Y. Cai, Q. Wu and T. Wang, *Chem.–Eur. J.*, 2015, **21**, 16754–16758.
- 149 A. Fukazawa, S. Suda, M. Taki, E. Yamaguchi, M. Grzybowski, Y. Sato, T. Higashiyama and S. Yamaguchi, *Chem. Commun.*, 2016, **52**, 1120–1123.
- 150 H. Ogasawara, M. Grzybowski, R. Hosokawa, Y. Sato, M. Taki and S. Yamaguchi, *Chem. Commun.*, 2018, **54**, 299–302.
- 151 P. Hindenberg, A. López-Andarias, F. Rominger, A. de Cózar and C. Romero-Nieto, *Chem.–Eur. J.*, 2017, **23**, 13919–13928.
- 152 E. Regulska and C. Romero-Nieto, *Dalton Trans.*, 2018, **47**, 10344–10359.
- 153 *Market Report Flame Retardants: Industry Analysis, Trends*, <https://www.ceresana.com/en/market-studies/chemicals/flame-retardants/>, accessed 12 August 2020.
- 154 A. Innes and J. Innes, in *Applied Plastics Engineering Handbook*, ed. M. Kutz, William Andrew Publishing, Oxford, 2011, pp. 469–485.
- 155 S. Shaw, *Rev. Environ. Health*, 2010, **25**(10), 261–305.
- 156 B. Schartel, *Materials*, 2010, **3**, 4710–4745.
- 157 E. D. Weil and S. V. Levchik, Phosphorus Flame Retardants, in *Kirk-Othmer Encyclopedia of Chemical Technology*, American Cancer Society, 2017, pp. 1–34.
- 158 F. Çelebi, O. Polat, L. Aras, G. Gündüz and I. M. Akhmedov, *J. Appl. Polym. Sci.*, 2004, **91**, 1314–1321.
- 159 M. Rakotomalala, S. Wagner and M. Döring, *Materials*, 2010, **3**, 4300–4327.
- 160 J. Zheng, Y. Yu, L. Zhang, X. Zhen and Y. Zhao, *Aust. J. Chem.*, 2014, **67**, 1688–1692.
- 161 J. Zheng, X. Li, Y. Yu, X. Feng and Y. Zhao, *J. Therm. Anal. Calorim.*, 2014, **117**, 319–324.
- 162 A. Blum, M. Behl, L. S. Birnbaum, M. L. Diamond, A. Phillips, V. Singla, N. S. Sipes, H. M. Stapleton and M. Venier, *Environ. Sci. Technol. Lett.*, 2019, **6**, 638–649.
- 163 J. Alongi, F. Cuttica and F. Carosio, *ACS Sustainable Chem. Eng.*, 2016, **4**, 3544–3551.
- 164 J. Alongi, A. Di Blasio, J. Milnes, G. Malucelli, S. Bourbigot, B. Kandola and G. Camino, *Polym. Degrad. Stab.*, 2015, **113**, 110–118.
- 165 J. Alongi, F. Carosio and G. Malucelli, *Polym. Degrad. Stab.*, 2014, **106**, 138–149.
- 166 M. Rahman and C. S. Brazel, *Prog. Polym. Sci.*, 2004, **29**, 1223–1248.
- 167 P. Walters, D. F. Cadogan and C. J. Howick, Plasticizers, in *Ullmann's Encyclopedia of Industrial Chemistry*, American Cancer Society, 2020, pp. 1–27.
- 168 M. Ehrich and B. S. Jortner, in *Hayes' Handbook of Pesticide Toxicology*, ed. R. Krieger, Academic Press, New York, 3rd edn, 2010, pp. 1479–1504.
- 169 D. N. Brooke, M. J. Crookes, P. Quarterman and J. Burns, *J. Environmental Risk Evaluation Report: 2-Ethylhexyl Diphenyl Phosphate (CAS No. 1241-94-7)*, UK Environment Agency, 2009.
- 170 M. Rahman and C. S. Brazel, *Polym. Degrad. Stab.*, 2006, **91**, 3371–3382.
- 171 A. E. Lemire, A. F. Janzen and K. Marat, *Inorg. Chim. Acta*, 1985, **110**, 237–241.
- 172 C. A. Blake, D. J. Crousse, C. F. Coleman, K. B. Brown and A. D. Kelmers, *Progress report: Further studies of the dialkylphosphoric acid extraction (DAPEX) process for uranium*, Oak Ridge National Laboratory, 1957.
- 173 J. R. Kumar, J.-S. Kim, J.-Y. Lee and H.-S. Yoon, *Sep. Purif. Rev.*, 2011, **40**, 77–125.
- 174 C. V. S. B. Rao, T. G. Srinivasan and P. R. V. Rao, *Solvent Extr. Ion Exch.*, 2010, **28**, 202–224.
- 175 D. Beltrami, G. Cote, H. Mokhtari, B. Courtaud, B. A. Moyer and A. Chagnes, *Chem. Rev.*, 2014, **114**, 12002–12023.
- 176 G. Koch, S. Träger, A. Max, W.-D. Krebs, W. Stoll, W. Heit, E. Warnecke, P. Brennecke and E. Merz, Nuclear Technology, 3. Fuel Cycle, in *Ullmann's Encyclopedia of Industrial Chemistry*, American Cancer Society, 2011.
- 177 J. S. Preston, P. M. Cole, W. M. Craig and A. M. Feather, *Hydrometallurgy*, 1996, **41**, 1–19.
- 178 X. Huang, J. Li, Z. Long, Y. Zhang, X. Xue and Z. Zhu, *J. Rare Earths*, 2008, **26**, 410–413.
- 179 K. Yoshizuka, H. Kosaka, T. Shinohara, K. Ohto and K. Inoue, *Bull. Chem. Soc. Jpn.*, 1996, **69**, 589–596.
- 180 X. Wang, W. Li, S. Meng and D. Li, *J. Chem. Technol. Biotechnol.*, 2006, **81**, 761–766.
- 181 E. Jorjani and M. Shahbazi, *Arabian J. Chem.*, 2016, **9**, S1532–S1539.
- 182 O. Klimchuk, L. Atamas, S. Miroshnichenko, V. Kalchenko, I. Smirnov, V. Babain, A. Varnek and G. Wipff, *J. Inclusion Phenom. Macrocyclic Chem.*, 2004, **49**, 47–56.
- 183 L. Chen, Y. Wang, X. Yuan, Y. Ren, N. Liu, L. Yuan and W. Feng, *Sep. Purif. Technol.*, 2018, **192**, 152–159.
- 184 K. Kiegiel, L. Steczek and G. Zakrzewska-Trznadel, *Application of Calixarenes as Macrocyclic Ligands for Uranium(vi)*, <https://www.hindawi.com/journals/jchem/2013/762819/>, accessed 14 August 2020.
- 185 N. Z. Burns, P. S. Baran and R. W. Hoffmann, *Angew. Chem., Int. Ed.*, 2009, **48**, 2854–2867.



- 186 M. B. Geeson and C. C. Cummins, *ACS Cent. Sci.*, 2020, **6**, 848–860.
- 187 J.-L. Montchamp, *Acc. Chem. Res.*, 2014, **47**, 77–87.
- 188 R. Engel, C–P bond formation using nucleophilic trivalent phosphorus reagents, in *Synthesis of Carbon–Phosphorus Bonds*, CRC Press, Boca Raton, 2003.
- 189 M. Caporali, M. Serrano-Ruiz and M. Peruzzini, in *Chemistry beyond Chlorine*, ed. P. Tundo, L.-N. He, E. Lokteva and C. Mota, Springer International Publishing, Cham, 2016, pp. 97–136.
- 190 J. E. Borger, A. W. Ehlers, J. C. Slootweg and K. Lammertsma, *Chem.–Eur. J.*, 2017, **23**, 11738–11746.
- 191 M. Scheer, G. Balázs and A. Seitz, *Chem. Rev.*, 2010, **110**, 4236–4256.
- 192 R. Engel, *Synthesis of Carbon–Phosphorus Bonds*, CRC Press, Boca Raton, 2003.
- 193 S. Hörold, N. Weferling and H. P. Breuer, *Eur. Pat.*, EP0969008A2, 2000.
- 194 N. Weferling, O. Stelzer, G. Kolbe, *Can. Pat.*, CA2311950A1, 1999.
- 195 M. M. Rauhut and A. M. Semsel, *J. Org. Chem.*, 1963, **28**, 471–473.
- 196 M. M. Rauhut and A. M. Semsel, *J. Org. Chem.*, 1963, **28**, 473–477.
- 197 D. G. Yakhvarov, Y. S. Ganushevich and O. G. Sinyashin, *Mendeleev Commun.*, 2007, **17**, 197–198.
- 198 K. M. Armstrong and P. Kilian, *Eur. J. Inorg. Chem.*, 2011, **2011**, 2138–2147.
- 199 D. H. R. Barton and J. Zhu, *J. Am. Chem. Soc.*, 1993, **115**, 2071–2072.
- 200 S. K. Ghosh, C. C. Cummins and J. A. Gladysz, *Org. Chem. Front.*, 2018, **5**, 3421–3429.
- 201 P. Barbaro, A. Ienco, C. Mealli, M. Peruzzini, O. J. Scherer, G. Schmitt, F. Vizza and G. Wolmershäuser, *Chem.–Eur. J.*, 2003, **9**, 5195–5210.
- 202 D. G. Yakhvarov, E. V. Gorbachuk and O. G. Sinyashin, *Eur. J. Inorg. Chem.*, 2013, **2013**, 4709–4726.
- 203 Y. H. Budnikova, D. G. Yakhvarov and O. G. Sinyashin, *J. Organomet. Chem.*, 2005, **690**, 2416–2425.
- 204 Y. H. Budnikova, S. A. Krasnov, T. V. Graznova, A. P. Tomilov, V. V. Turigin, I. M. Magdeev and O. G. Sinyashin, *Phosphorus, Sulfur Silicon Relat. Elem.*, 2008, **183**, 513–518.
- 205 M. Pecht and Y. Deng, *Microelectron. Reliab.*, 2006, **46**, 53–62.
- 206 U. Braun and B. Schartel, *Macromol. Chem. Phys.*, 2004, **205**, 2185–2196.
- 207 Y. Zhu, J. Ren, X. Zhang and D. Yang, *Nanoscale*, 2020, **12**, 13297–13310.
- 208 B. Trofimov, N. Gusarova and L. Brandsma, *Phosphorus, Sulfur Silicon Relat. Elem.*, 1996, **109**, 601–604.
- 209 N. K. Gusarova, S. I. Shaikhudinova, A. M. Reutskaya, N. I. Ivanova, A. A. Tatarinova and B. A. Trofimov, *Russ. Chem. Bull.*, 2000, **49**, 1320.
- 210 T. M. Wrodnigg, A. E. Stütz and S. G. Withers, *Tetrahedron Lett.*, 1990, **31**, 5463–5466.
- 211 M. C. J. M. V. Hooijdonk, G. Gerritsen and L. Brandsma, *Phosphorus, Sulfur Silicon Relat. Elem.*, 2000, **162**, 39–49.
- 212 L. Brandsma, S. Arbuzova, R.-J. D. Lang, N. Gusarova and B. Trofimov, *Phosphorus, Sulfur Silicon Relat. Elem.*, 1997, **126**, 125–128.
- 213 D. Albouy, G. Etemad-Moghadam and M. Koenig, *Eur. J. Org. Chem.*, 1999, **1999**, 861–868.
- 214 N. K. Gusarova, S. N. Arbuzova and B. A. Trofimov, *Pure Appl. Chem.*, 2012, **84**, 439–459.
- 215 B. A. Trofimov, S. N. Arbuzova and N. K. Gusarova, *Russ. Chem. Rev.*, 1999, **68**, 215.
- 216 F. Alonso, Y. Moglie, G. Radivoy and M. Yus, *Green Chem.*, 2012, **14**, 2699–2702.
- 217 M. M. Rauhut, H. A. Currier, A. M. Semsel and V. P. Wystrach, *J. Org. Chem.*, 1961, **26**, 5138–5145.
- 218 R. H. Hall and H. G. Khorana, *J. Am. Chem. Soc.*, 1954, **76**, 5056–5060.
- 219 M. Smith and H. G. Khorana, *J. Am. Chem. Soc.*, 1958, **80**, 1141–1145.
- 220 R. F. Long and A. L. Morrison, *J. Chem. Soc.*, 1954, 3854–3856.
- 221 E. Cherbuliez and J.-P. Leber, *Helv. Chim. Acta*, 1952, **35**, 644–664.
- 222 E. Cherbuliez and J. Rabinowitz, *Helv. Chim. Acta*, 1958, **41**, 1168–1175.
- 223 E. Cherbuliez and J. Rabinowitz, *Helv. Chim. Acta*, 1956, **39**, 1455–1461.
- 224 F. D. Popp and W. E. McEwen, *Chem. Rev.*, 1958, **58**, 321–401.
- 225 M. B. Geeson and C. C. Cummins, *Science*, 2018, **359**, 1383–1385.
- 226 D. Pham Minh, J. Ramaroson, A. Nzihou and P. Sharrock, *Ind. Eng. Chem. Res.*, 2012, **51**, 3851–3854.
- 227 C. J. Besecker, V. W. Day and W. G. Klemperer, *Organometallics*, 1985, **4**, 564–570.
- 228 M. B. Geeson, P. Ríos, W. J. Transue and C. C. Cummins, *J. Am. Chem. Soc.*, 2019, **141**, 6375–6384.
- 229 R. Gilmour, Phosphorus Compounds, in *Kirk-Othmer Encyclopedia of Chemical Technology*, American Cancer Society, 2019, pp. 1–57.
- 230 S. Deprère and J.-L. Montchamp, *J. Organomet. Chem.*, 2002, **643–644**, 154–163.
- 231 E. E. Nifant'ev and L. P. Levitan, *J. Gen. Chem. USSR*, 1965, **35**, 762.
- 232 J.-L. Montchamp, *Phosphorus, Sulfur Silicon Relat. Elem.*, 2013, **188**, 66–75.
- 233 J.-L. Montchamp, *Pure Appl. Chem.*, 2019, **91**, 113–120.
- 234 J.-L. Montchamp, *J. Organomet. Chem.*, 2005, **690**, 2388–2406.
- 235 H. C. Fisher, L. Prost and J.-L. Montchamp, *Eur. J. Org. Chem.*, 2013, **2013**, 7973–7978.

