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## Hydrogen spillover in Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy decorated Ni<sub>3</sub>S<sub>2</sub> enabling pH-universal electrocatalytic hydrogen evolution†

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As a new type of clean energy, hydrogen receives extensive attention due to its zero-pollution characteristics. Hydrogen produced by electrocatalysis is extensively studied due to its convenient properties. However, most catalysts for catalyzing the hydrogen evolution reaction are noble metals, which hinders their commercialization. Herein, we design a very effective electrocatalyst of Ni<sub>3</sub>S<sub>2</sub> decorated with a small amount of Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy supported on Ni foam (Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF) to enhance its hydrogen evolution reaction kinetics by exploiting the hydrogen spillover effect between Pt<sub>5</sub>Ru<sub>1</sub> and Ni<sub>3</sub>S<sub>2</sub>. The designed Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF exhibits excellent HER performance under both acidic and alkaline conditions, achieving a current density of 10 mA cm<sup>-2</sup> at only 13 mV and 42 mV, respectively.

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### Broader context

As a clean energy source with zero pollution, hydrogen has been paid more and more attention. Electrocatalytic hydrogen production is an effective and rapid synthetic method. Most electrocatalytic hydrogen production relies on noble metal catalysts, which increases the cost of hydrogen synthesis. At present, many non-noble metal catalysts have been reported for the hydrogen evolution reaction, but most of them have poor performance. In this work, we designed a Ni<sub>3</sub>S<sub>2</sub> catalyst, which was decorated with a small amount of Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy. This structure can effectively reduce the hydrogen evolution barrier by the hydrogen spillover effect. As a highly efficient hydrogen evolution catalyst, the designed Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy decorated Ni<sub>3</sub>S<sub>2</sub> (Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF) can significantly reduce the hydrogen evolution overpotential. This work could provide theoretical guidance and experimental evidence for designing high catalytic activity and high stability hydrogen evolution electrocatalysts and give fundamental insight into the potential construction of advanced electrocatalysts. We are sure that our work is very valuable to the readers of this journal and will attract a lot of attention.

## 1. Introduction

As the world continues to transition toward carbon emission-free energy technologies, it is necessary for us to reduce the carbon emissions of the chemical production industry.<sup>1</sup> Hydrogen, featured by high gravimetric energy density and zero-emission, holds great potential as a sustainable alternative to fossil fuels.<sup>2</sup> Electrocatalytic water splitting is an efficient method for producing high-purity hydrogen, but this method is far from cost-competitive with commercial steam reforming

for hydrogen production.<sup>3</sup> At present, people are devoted to finding electrocatalysts with high catalytic performance and long life in extreme acid–base environments.<sup>4,5</sup> In general, catalyst design for electrocatalytic hydrogen evolution reactions is guided by the classical volcano theory, in which the free energy of hydrogen adsorption on the catalyst surface is an important indicator.<sup>6</sup> The theory points out that the adsorption energy for protons is neither too strong nor too weak, which should favor the protons gaining electrons at the active sites and easily leaving the surface to form hydrogen gas.<sup>7</sup> The elements at the top of the volcano map have thus been a hot research topic. Platinum, a representative of the precious metals, has become the target of choice for the commercial electrocatalytic production of hydrogen.<sup>8</sup> However, its high price and weak hydrogen evolution performance under alkaline conditions hinder its further application.<sup>9</sup> Therefore, the development of electrocatalysts with no noble metal or low noble metal content is the focus of current research.

Various terrestrially abundant and inexpensive transition metal compounds have been explored, among which sulfides have

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been extensively studied due to their novel physicochemical properties.<sup>10</sup> There are numerous reports on the role of sulfides in electrocatalytic hydrogen evolution, among which the Ni<sub>3</sub>S<sub>2</sub> catalyst stands out due to its unique electronic structure.<sup>11–15</sup> However, recent studies have shown that the inherent activity of Ni<sub>3</sub>S<sub>2</sub> is insufficient, and its intrinsic catalytic activity is at least two orders of magnitude worse than that of noble metals,<sup>16</sup> which may be related to the unfavorable free energy of hydrogen adsorption on the surface of Ni<sub>3</sub>S<sub>2</sub>. Many methods have been utilized to modify the Ni<sub>3</sub>S<sub>2</sub> catalyst, including doping, introduction of vacancy defects, and interfacial engineering, but simply modifying the electronic and geometric structures of the Ni<sub>3</sub>S<sub>2</sub> catalyst cannot effectively compensate for its intrinsically poor activity. Inspired by the phenomenon of hydrogen spillover induced by strong metal–support interactions during thermal hydrogenation, hydrogen spillover has recently emerged as a research frontier in HER electrocatalysts.<sup>17,18</sup> This strategy utilizes the difference in the free energy of hydrogen adsorption between metals and supports to kinetically facilitate the adsorption and desorption of protons.<sup>19,20</sup> The binary ‘metal–support’ system with the hydrogen spillover effect has great potential to boost the performance of the Ni<sub>3</sub>S<sub>2</sub> catalyst with less metal usage.

Herein, we designed a binary ‘metal–support’ electrocatalyst based on Ni<sub>3</sub>S<sub>2</sub> decorated with Pt and Ru nanoalloys on Ni foam (NF). By introducing a small amount of noble metal elements, the hydrogen evolution performance of Ni<sub>3</sub>S<sub>2</sub>@NF in acidic and alkaline solution can be greatly improved. Meanwhile, compared with Pt nanoparticles, the work function difference ( $\Delta\phi$ ) between the Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy and Ni<sub>3</sub>S<sub>2</sub> substrate is smaller, which effectively prevents charge accumulation at the interface and enhances the hydrogen overflow process. The unique structure of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF provides a new idea for designing high catalytic activity and high stability hydrogen evolution electrocatalysts.

## 2. Results and discussion

### Microstructure characterization

The Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF catalyst was prepared by a one-step hydrothermal method, followed by a chemical reduction reaction. Specifically, Ni<sub>3</sub>S<sub>2</sub> was synthesized *in situ* on a nickel foam substrate by a hydrothermal method, and then H<sub>2</sub>PtCl<sub>6</sub> and RuCl<sub>3</sub> were reduced at room temperature to obtain the Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF heterostructure catalyst. The detailed process is illustrated in the Experimental section. X-ray diffraction (XRD) characterization was employed to identify the crystal structure of the as-prepared Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF sample. As presented in Fig. 1a, two dominant peaks appearing at  $2\theta$  of 44.5° and 51.8° are assigned to the pristine Ni foam substrate (PDF#04-0850).<sup>21</sup> Additionally, other diffraction peaks are well matched with hexagonal heazlewoodite Ni<sub>3</sub>S<sub>2</sub> (PDF#44-1418).<sup>22</sup> Due to the small size and low loading of the catalyst, no XRD peaks for the supported metal are clearly observed. The weight ratio of supported metal catalysts of Pt to Ru was determined to be 5 : 1 by inductively-coupled plasma mass spectrometry (ICP-MS,

Table S1, ESI†). Scanning electron microscopy (SEM) and transmission electron microscopy (TEM) were carried out to investigate the morphology and microstructure of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF. The SEM images (Fig. S1, ESI†) show that Ni<sub>3</sub>S<sub>2</sub> is densely loaded on the surface of nickel foam, and Pt<sub>5</sub>Ru<sub>1</sub> nanoalloys are grown on Ni<sub>3</sub>S<sub>2</sub>. In contrast, the surface of the Ni foam is smoother (shown in Fig. S2a, ESI†) and the surface graininess was increased after loading Ni<sub>3</sub>S<sub>2</sub>, as shown in Fig. S2b (ESI†). The TEM image (Fig. 1b) also shows that nanoparticles with an average size of 5 nm are uniformly decorated on the supported submicron particles. The corresponding FFT patterns (insets in Fig. 1b) reveal that the particles really are Pt<sub>5</sub>Ru<sub>1</sub> (*Fm* $\bar{3}$ *m*), and Ni<sub>3</sub>S<sub>2</sub> (R32). The interplanar spacings of 0.287 and 0.225 nm correspond to the (110) facet of Ni<sub>3</sub>S<sub>2</sub> and (111) facet of Pt, respectively. Owing to the close interplanar spacing between Ru and Pt, Ru atoms can fill in the Pt lattice to form nanoalloys in a ratio of 5 : 1, which is proved by line scanning (Fig. S3, ESI†). Additionally, the STEM image (Fig. 1c) and its corresponding elemental mappings prove that Pt<sub>5</sub>Ru<sub>1</sub> nanoalloys are evenly distributed on the surface of the Ni<sub>3</sub>S<sub>2</sub> substrate. The chemical compositions and valence states of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF were examined by X-ray photoelectron spectroscopy (XPS). As shown in Fig. 1d–g, the XPS spectra of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF testify the presence of elemental Ni, S, Pt and Ru, which is consistent with the EDX results. As for the high-resolution Ni 2p spectrum of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF in Fig. 1d, two visible peaks located at binding energies of 873.6 and 855.7 eV are assigned to the Ni 2p<sub>1/2</sub> and Ni 2p<sub>3/2</sub> orbital of Ni<sup>2+</sup>, respectively.<sup>23,24</sup> Compared with Ni<sub>3</sub>S<sub>2</sub>@NF (Fig. S4, ESI†), the peaks of Ni in Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF shift to higher energy, indicating that there is interfacial electron transfer between the Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy and Ni<sub>3</sub>S<sub>2</sub>, and more electrons flow to the Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy.<sup>25</sup> No obvious change has been observed in the peaks of S (Fig. 1e), indicating that the electron transfer mainly occurs between the metal elements. Compared with Pt foil, the Pt in Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF shows a 0.5 eV negative shift in Fig. 1f.<sup>15</sup> And a slight negative shift of Ru in Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF is also observed in Fig. 1g compared to Ru foil. The electron transfer process between Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy and Ni<sub>3</sub>S<sub>2</sub> can also be confirmed by the work function calculation below.

### Acidic and alkaline HER performance

A typical three-electrode system was used to evaluate the HER performance of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF in 1.0 M KOH and 0.1 M HClO<sub>4</sub> solutions. For comparison, we also synthesized control samples of Ni<sub>3</sub>S<sub>2</sub>/Pt@NF, Ni<sub>3</sub>S<sub>2</sub>/Ru@NF and commercial PtC@NF catalysts with similar noble metal loadings. The performance of the nickel foam substrate was also tested to prove that our catalytic activity mainly comes from the Ni<sub>3</sub>S<sub>2</sub> and nanoalloys. The polarization curves of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF and the control samples are displayed in Fig. 2a, revealing that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF can reach a current density of 10 mA cm<sup>−2</sup> with only a low overpotential of 13 mV in 1.0 M KOH, exceeding most Ni<sub>3</sub>S<sub>2</sub>-based catalysts.<sup>15,26,27</sup> Fig. 2b compares the overpotential values of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF and the control group at current densities of 10, 100 and 500 mA cm<sup>−2</sup>. The results show that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF





Fig. 1 Structural and microstructure characterizations of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$ . (a) XRD pattern, (b) HRTEM image and the corresponding FFT patterns, (c) STEM image and the corresponding EDX elemental mapping images, and XPS spectra of (d) Ni 2p, (e) S 2p, (f) Pt 4f, and (g) Ru 3d.

has the smallest overpotential values at different current densities, and the performance is still better than that of commercial Pt/C. Electrochemically active surface area (ECSA) is often used to evaluate the actual catalytic performance of electrode materials. Since there is a linear relationship between ECSA and electric double-layer capacitance ( $C_{dl}$ ), we utilized cyclic voltammetry (CV) to measure the  $C_{dl}$  of the catalyst at different scan rates, as shown in Fig. 2c. It can be observed that  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  ( $22.7 \text{ mF cm}^{-2}$ ) has a  $C_{dl}$  close to that of Pt/C ( $26.9 \text{ mF cm}^{-2}$ ), and significantly higher than other control samples. The polarization curves of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  and control samples based on ECSA-normalization toward the HER in 1.0 M KOH are displayed in Fig. S5b (ESI<sup>†</sup>). The results show that  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  also has the smallest overpotential values at different current densities. Fig. 2d tests the Tafel slopes of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  and the control samples. It is found that  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  has a smaller Tafel slope, indicating a faster kinetic process at low potentials. The electrochemical impedance spectroscopy (EIS) of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  and the control samples was performed and simulated by a double-parallel equivalent circuit model (Fig. 2e and Table S2, ESI<sup>†</sup>). The first parallel component ( $T$  and  $R_1$ ) reflects the charge transfer kinetics,

where  $T$  is related to the electric double layer capacitance and  $R_1$  represents the catalytic charge transfer resistance.  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  has the smallest resistance after fitting, indicating its faster charge transfer process. The second parallel component ( $C_\psi$  and  $R_2$ ) reflects the hydrogen adsorption behavior on the catalyst surface, where  $C_\psi$  and  $R_2$  represent the hydrogen adsorption pseudo-capacitance and resistance, respectively. It can be observed from the fitting results in Table S2 (ESI<sup>†</sup>) that  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  has a larger  $C_\psi$  and a smaller  $R_2$ , indicating that its unique ‘metal-support’ structure is more conducive to the binding of protons and the desorption of hydrogen on the surface. We also tested the long-term stability by chronopotentiometry measurement (Fig. 2f). At a constant current of  $10 \text{ mA cm}^{-2}$ , there is basically no obvious attenuation of the overpotential after 40 h for the  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$ , which formed a clear contrast with the commercial Pt/C catalyst after 2 h.

Likewise, the HER performance of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  in 0.1 M  $\text{HClO}_4$  was also investigated. From the polarization curves in Fig. 3a and the comparison of the overpotential at different current densities in Fig. 3b,  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@\text{NF}$  has similar overpotential to commercial Pt/C at a current density of





Fig. 2 Electrocatalytic HER performance in 1.0 M KOH solution. (a) Polarization curves of the electrocatalysts for the HER (with 95% iR compensation), (b) the resultant overpotentials at current densities of 10, 100 and 500 mA cm<sup>-2</sup> for the as-prepared electrocatalysts, (c) calculated electrochemical double-layer capacitances, (d) the corresponding Tafel plots, (e) Nyquist plots under open-circuit voltage, and (f) long-term stability test at current densities of 10 mA cm<sup>-2</sup>.

10 mA cm<sup>-2</sup>. However, with the increase of the current density, the HER performance of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF gradually exceeds that of commercial Pt/C, and the overpotential of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF is even only half that of commercial Pt/C at a large

current density of 100 mA cm<sup>-2</sup>. This may be related to the extremely fast kinetic process of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF in acidic solution, which can also be demonstrated by the Tafel slope in Fig. 3c. Likewise, the polarization curves of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF

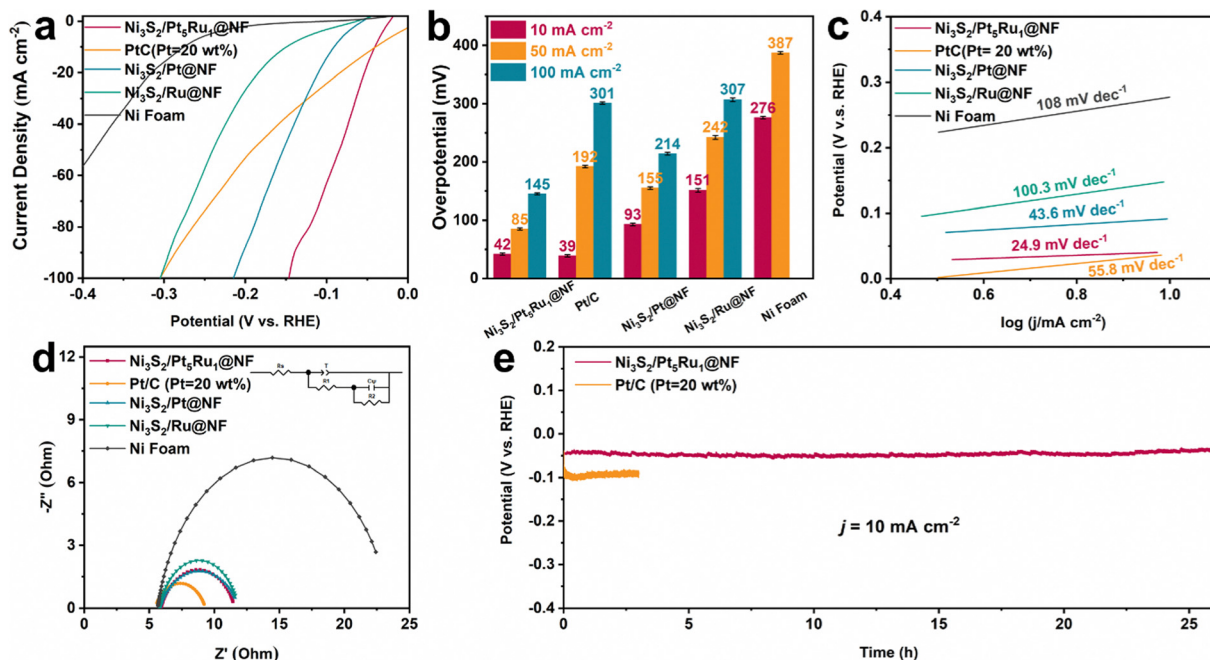


Fig. 3 Electrocatalytic HER performance in 0.1 M HClO<sub>4</sub> solution. (a) Polarization curves of the electrocatalysts for the HER (with 95% iR compensation), (b) the resultant overpotentials at current densities of 10, 50 and 100 mA cm<sup>-2</sup> for the as-prepared electrocatalysts, (c) the corresponding Tafel plots, (d) Nyquist plots under open-circuit voltage, and (e) long-term stability test at current densities of 10 mA cm<sup>-2</sup>.



and the control samples based on ECSA-normalization toward the HER in 0.1 M HClO<sub>4</sub> show a similar regularity (Fig. S5a, ESI<sup>†</sup>). The EIS of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF and control samples under acidic conditions were also fitted, as shown in Fig. 3d and Table S3 (ESI<sup>†</sup>). Like the results in 1.0 M KOH, the unique structural design of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF provides a good conductive network for the rapid transfer of protons and desorption of hydrogen. More excitingly, Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF exhibits excellent durability in 0.1 M HClO<sub>4</sub> with almost no decay of the overpotential at a current density of 10 mA cm<sup>-2</sup> (Fig. 3e), indicating that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF can exist stably for a long time in acidic solution. Compared with other nickel-based chalcogenide catalysts, the overpotential of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF is in a leading position, as shown in Table S4 (ESI<sup>†</sup>).

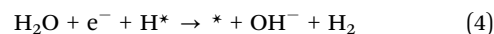
### Experimental evidence for hydrogen spillover

The hydrogen spillover process is widespread on catalyst surfaces with at least two active sites, usually from sites with more negative ΔG<sub>H</sub> to those with more positive ΔG<sub>H</sub>. Generally, the HER involves a two-electron process, and the reaction pathway is as follows:

Acidic solution:



Alkaline solution:



In acidic solution, H<sup>+</sup> is preferentially adsorbed on the surface of the active site, taking electrons from the bulk to form protons. Proton and proton combine to form H<sub>2</sub> leaving the active site. For alkaline solutions, the first step is often the dissociation of H<sub>2</sub>O in the active site. Herein, if it is necessary to prove the occurrence of the hydrogen spillover process, two points must be understood: (i) the intrinsic activity of the active site, and (ii) the change of the surface adsorption state. Therefore, we first tested the turnover frequency (TOF) of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF. It can be seen from Fig. 4a that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF has the highest TOF value compared to the control samples, indicating that the active sites of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF have high intrinsic activity. The kinetic isotope effect (KIE) is normally used to estimate the involvement of proton transfer in the rate-determining step of an electrocatalytic reaction.<sup>29</sup> As shown in Fig. 4b, the KIE value of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF is over 1, implying that protons are involved in the HER reaction.<sup>30</sup> Compared with control samples, Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF has the smallest KIE value, which indicates that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF has an extremely fast proton transfer kinetic process.<sup>31</sup> In order to further confirm the active sites of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF for the HER, poisoning tests were performed by adding thiocyanate (SCN<sup>-</sup>) and tetramethylammonium cation (TMA<sup>+</sup>) ions into acidic and alkaline solutions.<sup>32</sup> It has been reported that TMA<sup>+</sup> has adsorption properties for O-group elements.<sup>33</sup> In Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF, it can be adsorbed on the S site to block the path of hydrogen



Fig. 4 Mechanism study of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF. (a) The potential-dependent TOF curves for Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF, Pt/C, Ni<sub>3</sub>S<sub>2</sub>/Pt@NF and Ni<sub>3</sub>S<sub>2</sub>/Ru@NF, (b) calculated KIE values ( $J_{\text{H}_2\text{O}}/J_{\text{D}_2\text{O}}$ ) from Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF, Pt/C, Ni<sub>3</sub>S<sub>2</sub>/Pt@NF, Ni<sub>3</sub>S<sub>2</sub>/Ru@NF and Ni foam, (c) polarization curves for Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF before and after the addition of SCN<sup>-</sup> and TMA<sup>+</sup> ions in N<sub>2</sub>-saturated 1.0 M KOH, and (d) *operando* Raman spectra of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF in 1.0 M KOH under HER potential.



overflow. As a commonly used ion in poisoning experiments,  $\text{SCN}^-$  can be adsorbed on the surface of the  $\text{Pt}_5\text{Ru}_1$  alloy to inactivate it. As shown in Fig. 4c, the addition of  $\text{SCN}^-$  and  $\text{TMA}^+$  affects the HER response to varying degrees.  $\text{TMA}^+$  blocks the path for hydrogen spillover, making the transfer of protons between active sites difficult, and  $\text{SCN}^-$  adsorption on the surface of  $\text{Pt}_5\text{Ru}_1$  alloy poisons the catalyst, dropping the activity of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@NF$  almost to zero. The above results indicate that both  $\text{Ni}_3\text{S}_2$  and  $\text{Pt}_5\text{Ru}_1$  are involved in the HER as active sites. By the extent of the decrease in activity in Fig. 4c, we predict that  $\text{Pt}_5\text{Ru}_1$  preferentially undergoes the Volmer step to generate protons due to its higher activity.

Although the above electrochemical methods can reveal the existence of active sites and the strength of activity in  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@NF$ , the real hydrogen spillover process is hard to identify. The hydrogen spillover process itself is dynamic, and transient imaging technology with ultra-high spatial

resolution and nanometer size should be used to monitor the hydrogen spillover process in real time.<sup>34</sup> Unfortunately, there is no available technology to visually observe the hydrogen spillover process. Considering that the catalytic reaction generally occurs at the surface interface, a surface-sensitive technique can be used to infer the occurrence of the hydrogen spillover process. Herein, we attempt to instantaneously record the HER of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@NF$  by operando Raman spectroscopy, and speculate the hydrogen spillover process *via* bonding changes on its surface.<sup>35,36</sup> The Raman spectra of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@NF$  and control samples were thus studied to obtain the information of surface bonding. As shown in Fig. S6a (ESI<sup>†</sup>), Raman signals appeared at 192, 224, 305, 326 and 353  $\text{cm}^{-1}$  for  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1@NF$  and the control samples correspond to the E4, E3, E2, A1 and E1 vibrational states of  $\text{Ni}_3\text{S}_2$ , respectively, indicating that the  $\text{Ni}_3\text{S}_2$  is abundant in the samples.<sup>27</sup> The signals of Pt and Ru are not detected, which may be related to the insensitivity of Raman spectroscopy



Fig. 5 DFT calculation, (a) calculated PDOS of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1$ ,  $\text{Ni}_3\text{S}_2$ , Ru foil and Pt foil with aligned Fermi level  $E_F$ , (b) the d-p orbital alignment of the surface Ru, Pt, S and Ni sites for  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1$ , (c) the local electron density distribution of the  $\text{Pt}_5\text{Ru}_1$  cluster in  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1$ , (d) calculated d-band center ( $E_d$ ) of  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1$ ,  $\text{Ni}_3\text{S}_2$ , Ru foil and Pt foil, (e) calculated free energy diagram for the HER on  $\text{Ni}_3\text{S}_2/\text{Pt}_5\text{Ru}_1$  paradigm and  $\text{Ni}_3\text{S}_2/\text{Pt}_6$  benchmark, and (f) the optimized  $\text{H}^*$  adsorption structures at various sites.



to metal elements. Fig. S6b (ESI<sup>†</sup>) compares the signals of 1.0 M KOH and Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF in 1.0 M KOH, indicating that the new peaks appearing in Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF are not the signals of KOH. Fig. S6c (ESI<sup>†</sup>) compares the spectral information of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF before and after adding 1.0 M KOH. After adding 1.0 M KOH, three new peaks appeared at 468, 549 and 1059 cm<sup>-1</sup> correspond to Ni–O and two Ni–S bonds, respectively. A broad peak appears at 1635 cm<sup>-1</sup>, associated with the H–OH bending vibration of adsorbed water. The above results indicate that after the addition of 1.0 M KOH, OH<sup>-</sup> and H<sub>2</sub>O in the solution will be adsorbed on the surface of Ni sites. Therefore, the hydrogen spillover process can be indirectly observed by change of the Ni–O bond on the surface of the Ni site at 468 cm<sup>-1</sup>. As shown in Fig. 4d, the peak intensity of the Ni–O bond gradually weakens as the potential increases. This is because with the increase of the potential, the hydrogen overflow process of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF is continuously enhanced, and more and more protons are generated on the surface of the Pt<sub>5</sub>Ru<sub>1</sub> alloy and overflow to the Ni sites in the Ni<sub>3</sub>S<sub>2</sub> to form H<sub>2</sub>. The adsorption state on the Ni surface gradually changed from Ni–O bonds to the adsorption of protons at Ni sites. This result is quite consistent with the hydrogen spillover process from the Pt<sub>5</sub>Ru<sub>1</sub> alloy to Ni<sub>3</sub>S<sub>2</sub> as we conceived.

### DFT calculations for hydrogen spillover

Density functional theory (DFT) calculations were further performed to gain insights into the hydrogen spillover process. Combining XRD and HRTEM results, we choose the (110) facet of Ni<sub>3</sub>S<sub>2</sub> as the substrate, and Pt<sub>5</sub>Ru<sub>1</sub> clusters on it as the nanoalloy.<sup>37</sup> Details of the simulation models and DFT calculations are illustrated in the ESI.<sup>†</sup> Fig. 5a calculates the DOS and PDOS of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>, Ni<sub>3</sub>S<sub>2</sub>, Ru foil and Pt foil. The results show that Pt<sub>5</sub>Ru<sub>1</sub> increases the occupied state of Ni<sub>3</sub>S<sub>2</sub> near the Fermi level, which means that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF can more easily obtain electrons from the bulk for the HER.<sup>38</sup> Fig. 5b calculates the PDOS of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub> from the angle of Pt, Ru, S and Ni elements. The 1s orbital of H and the 2pπ orbital of H<sub>2</sub>O overlap well with Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>. This means that H<sub>2</sub>O or protons can make good contact with various elements at the interface of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF and transfer electrons during the reaction, which also explains the lower charge transfer resistance of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF in Fig. 2e and 3d. To intuitively describe the interaction between Ni<sub>3</sub>S<sub>2</sub> substrate and Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy, we calculated the charge density difference of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>. As shown in Fig. 5c, the Pt<sub>5</sub>Ru<sub>1</sub> nanoalloy and Ni<sub>3</sub>S<sub>2</sub> surface have good electron transfer, which means that Pt<sub>5</sub>Ru<sub>1</sub> can be stably supported on Ni<sub>3</sub>S<sub>2</sub> without being detached from the surface.<sup>26</sup> The d-band centers of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>, Ni<sub>3</sub>S<sub>2</sub>, Pt foil and Ru foil were calculated by DOS, as shown in Fig. 5d. The lower the d-band center, the lower the position of the antibonding orbital and the lower the adsorption energy. According to previous literature reports, Pt and Ru have the most excellent adsorption energy, allowing protons to react and forming H<sub>2</sub> off the surface.<sup>39,40</sup> Compared with Ni<sub>3</sub>S<sub>2</sub>, Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub> has a d-band center closer to Pt and Ru, so protons are more likely to be adsorbed and desorbed on the surface of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>. The above calculations indicate that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF itself

has good electrical conductivity and moderate hydrogen adsorption energy, and is an excellent catalyst for the HER.

The above evidence, however, is not enough to prove that Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF can generate hydrogen spillover. In theory, two conditions are required to produce hydrogen spillover: (i) There is a difference in the free energy of hydrogen evolution ( $G_H$ ) between the active sites. The ions preferentially undergo the Volmer step to form protons at the more negative  $\Delta G_H$  sites, and the protons then transfer to the more positive  $\Delta G_H$  sites to undergo the Tafel step to form the H<sub>2</sub>. (ii) The transfer between active sites should have a smaller migration energy barrier to reduce obstruction of hydrogen spillover. For the convenience of comparison, we chose Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>6</sub> as the control sample, and took the HER reaction mechanism under acidic conditions as the overall design framework. The  $\Delta G_H$  of Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub> and Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>6</sub> was calculated in an attempt to construct a hydrogen spillover path. As shown in Fig. 5e and f, the optimized hydrogen spillover path reveals that the uniquely ‘metal–support’ structured Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub> has a suitable  $\Delta G_H$  compared with Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>6</sub>. H<sup>+</sup> preferentially adsorbs to site 1 of the Pt<sub>5</sub>Ru<sub>1</sub> alloy with a  $\Delta G_H$  of -0.43 eV, and gains electrons from the Pt<sub>5</sub>Ru<sub>1</sub> alloy site to form protons. It is then transferred to the interface of the alloy and substrate through site 2 and site 3 of Pt<sub>5</sub>Ru<sub>1</sub>, and finally H<sub>2</sub> leaves the surface with a  $\Delta G_H$  of 0.09 eV at the Ni site of Ni<sub>3</sub>S<sub>2</sub>. On the other hand, it is well known that due to the difference in work function between the metal and the support, a Schottky barrier is inevitably formed, which traps protons at the interface and hinders hydrogen spillover. Therefore, reducing the work function difference ( $\Delta\phi$ ) as much as possible is an effective method to enhance the hydrogen spillover process. Here, we calculated  $\Delta\phi$  between Ni<sub>3</sub>S<sub>2</sub> and Pt (Pt<sub>5</sub>Ru<sub>1</sub>). The calculated results are shown in Fig. S7 and S8 (ESI<sup>†</sup>).  $\Delta\phi$  between Pt<sub>5</sub>Ru<sub>1</sub> and Ni<sub>3</sub>S<sub>2</sub> is 0.5678 eV, which is significantly smaller than 0.8442 eV between Pt and Ni<sub>3</sub>S<sub>2</sub>. That is to say, the Schottky barrier at the Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub> interface is smaller, and protons are not easily captured at the interface, which is favorable for the occurrence of the hydrogen spillover process.

## 3. Conclusions

In summary, through a simple two-step synthesis, we designed a ‘metal–support’ structured Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF catalyst with excellent catalytic performance under both acidic and basic conditions. Through the combination of experimental verification and theoretical calculation, the mechanism of the hydrogen spillover process is explained in detail. But as a new theory, the hydrogen spillover process itself needs more characterization methods to clarify its physical and chemical processes. As a successful case, Ni<sub>3</sub>S<sub>2</sub>/Pt<sub>5</sub>Ru<sub>1</sub>@NF provides strong proof for the development of this theory.

## Author contributions

Y. Y. and X. R. conceived and supervised the research. Y. Y. and Z. Y. designed the experiments. Z. Y. performed the



experiments and data analysis. Z. Y. conducted DFT calculations. All the authors discussed the results and commented on the paper.

## Conflicts of interest

There are no conflicts to declare.

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