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Hydrobiogechemical interactions in the hyporheic zone of a sulfate-impacted, freshwater stream and riparian wetland ecosystem†‡

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Coupled abiotic and biotic processes in the hyporheic zone, where surface water and groundwater mix, play a critical role in the biogeochemical cycling of carbon, nutrients, and trace elements in streams and wetlands. Dynamic hydrologic conditions and anthropogenic pollution can impact redox gradients and biogeochemical response, although few studies examine the resulting hydrobiogeochemical interactions generated within the hyporheic zone. This study examines the effect of hyporheic flux dynamics and anthropogenic sulfate loading on the biogeochemistry of a riparian wetland and stream system. The hydrologic gradient as well as sediment, surface water, and porewater geochemistry chemistry was characterized at multiple points throughout the 2017 spring-summer-fall season at a sulfate-impacted stream flanked by wetlands in northern Minnesota. Results show that organic-rich sediments largely buffer the geochemical responses to brief or low magnitude changes in hydrologic gradient, but sustained or higher magnitude fluxes may variably alter the redox regime and, ultimately, the environmental geochemistry. This has implications for a changing climate that is expected to dramatically alter the hydrological cycle. Further, increased sulfate loading and dissolved or adsorbed ferric iron complexes in the hyporheic zone may induce a cryptic sulfur cycle linked to iron and carbon cycling, as indicated by the abundance of intermediate valence sulfur compounds (*e.g.*, polysulfide, elemental sulfur, thiosulfate) throughout the anoxic wetland and stream-channel sediment column. The observed deviation from a classical redox tower coupled with potential changes in hydraulic gradient in these organic-rich wetland and stream hyporheic zones has implications for nutrient, trace element, and greenhouse gas fluxes into surface water and groundwater, ultimately influencing water quality and global climate.

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Environmental significance

The hyporheic zone below freshwater streams and wetlands, where mixing of surface water and groundwater occurs, plays a critical role in water quality and ecosystem health. Our understanding of the interactions between dynamic water fluxes and environmental geochemistry is limited. In this study, we examine fluctuating water flow through a sulfate-impacted riparian wetland and stream system to assess the geochemical impacts on the subsurface. We show that sediment and porewater composition can buffer the impacts of brief or small magnitude hydrologic changes, but more substantial flux changes can lead to measurable geochemistry outcomes. Further, we show that sulfur cycling is coupled to iron and methane cycling, which has implications for fluxes of carbon, nutrients, and contaminants in the environment.

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Introduction

The hyporheic zone, the layer of sediment where oxygen-rich surface water mixes with nutrient-rich shallow groundwater, plays a critical role in controlling water quality and global ecosystem health.[1] In the hyporheic zones of streams and associated riparian wetlands, high concentrations of organic matter[2,3] combined with saturated and typically flooded sediment beds promote the fast consumption of oxygen in these aquatic ecotones.[4–6] These conditions set up steep oxidation-reduction (redox) gradients that promote biogeochemical activity and organic matter turnover – it is estimated that up to 96% of riparian-river ecosystem respiration occurs within the hyporheic zone.[7] Water flow is also spatially and temporally heterogeneous in the hyporheic zone[8–10] which influences solute transport and further drives redox zone fluctuations that alter the geochemistry as well as the microbial community composition, function and activity.[3,8,11–18] These dynamic hydro-biogeochemical interactions drive the formation and fluxes of major nutrients,[4,19,20] redox active metals,[11,21] contaminants,[22–24] and greenhouse gases,[15,25] having implications that reach far beyond the ecotone boundaries.

High organic matter turnover in the hyporheic zone is fueled by microbial metabolic processes that regulate the biogeochemical cycling of carbon and other nutrients.[26] The "thermodynamic" redox ladder concept predicts the sequence of redox reactions that drive these metabolic processes where the more energetically favorable terminal electron acceptor (TEA) will be consumed before the next favorable electron acceptor.[27] For example, dissimilatory nitrate reduction should not occur until all oxygen (the more favorable electron acceptor) is depleted. This can lead to vertical stratification of the redox zones in some environments (*e.g.*, marine sediments), forming a redox gradient with more oxidized conditions at the surface due to transport of oxygen from the surface waters and more reduced conditions deeper in the subsurface.[28]

Typically in the anoxic portions of the hyporheic zone, microbial nitrate and ferric iron (Fe(III)) reduction as well as methanogenesis are dominant processes, resulting in elevated nitrous oxide, dinitrogen, ferrous iron (Fe(II)), and methane (CH_4) concentrations in interstitial porewaters.[15,21,29] Depending on the hydrologic conditions, these compounds can be transported to the surface water and atmosphere.[30–32] With relatively low sulfate concentrations, dissimilatory sulfate reduction has traditionally not been expected to be an important process in freshwater environments based on the hierarchical redox ladder. This paradigm, however, has been changing as several studies[33–35] identified fast microbial sulfate reduction rates sustained by cryptic sulfur (S) cycling coupled to Fe cycling in anoxic environments. These cryptic reactions involve oxidation of sulfide (produced by sulfate reduction) to intermediate-valence S forms (*e.g.*, thiosulfate, S(0), polysulfide), and disproportionation of these S intermediates can yield more sulfate to sustain the cycle. Sulfate is abundant in many industrial and agricultural wastewaters,[36–38] so cryptic S cycling may be increasingly recognized as an important driver in organic matter decomposition in many freshwater environments impacted by wastewaters. Further, cryptic S cycling has implications for methane fluxes to the atmosphere because methane can be anaerobically oxidized (consumed) coupled to sulfate reduction pathways in the subsurface.[39–41] In natural environments that are physically and chemically heterogeneous, like hyporheic zones, the redox ladder concept is being further challenged. Studies show that several redox reactions may co-occur or deviate from the predicted hierarchical order due to factors such as geochemical microniches,[42–44] hydrologic conditions,[8,10] and redox phase metastability.[45]

Hydrobiogeochemical interactions in hyporheic zones are complex and dynamic, and our understanding of these interactions and their environmental impacts is rapidly evolving. It is expected that changes in hydrologic flux direction or magnitude in the hyporheic zone would have a measurable impact, perhaps reflected in microsites ("hot spots") and/or for short times ("hot moments"), on the biogeochemistry of the sediments and porewaters.[10,42,44,46–50] The change from a gaining stream (upward groundwater flux) to a losing stream (downward surface water flux), for example, could alter the redox gradient by recharging the system with thermodynamically favorable TEAs (*e.g.*, oxygen, sulfate, iron oxides, *etc.*). This could lead to changes in biogeochemical reactions and rates that would be reflected in the sediment and porewater geochemistry.[11–18] The oxidation of aqueous Fe(II) and Fe(II)-bearing sulfides to form Fe(III)-oxides was previously observed, for example, in mangrove sediments due to tidal fluctuations and fluxes of oxygenated water into highly reducing sediments.[51] These changes, however, might be temporally and spatially variable for the different chemical species. Many questions remain about the impact of pollution and fluctuating hydraulic flux direction on the biogeochemistry of the system, such as the rates, spatial and temporal heterogeneity, and overall extent of geochemical change during or after changes in hydraulic regime.

To shed light on these poorly resolved processes, reactive transport modeling was conducted in previous studies by Ng *et al.*,[52,53] of a sulfate-impacted freshwater stream and associated riparian wetlands. Calibrated simulation results showed that sulfate reduction rates exceeded thermodynamically favored Fe reduction. Cryptic S cycling was proposed as a driving mechanism for the high rates, but key intermediate-valence S forms were not yet confirmed and thus not included in the modeled reactions. Interestingly, with the incorporation of only complete reduction and aerobic reoxidation reactions between sulfate to sulfide, the model could not properly reproduce the measurable concentrations of sulfate observed at depth – modeled sulfate concentrations always decreased to 0. This suggested that, in reality, anoxic oxidation of sulfide (not represented in the model) could be regenerating sulfate at depth. Additionally, the model presumed Fe(III) presence throughout the hyporheic zone, despite the absence of geochemical measurements to confirm this. If Fe(III) is indeed present throughout the hyporheic zone, this would suggest that anaerobic sulfide oxidation could be coupled to abiotic Fe(III) reduction as part of cryptic S cycling in this Fe-rich ecotone. To identify the activity and impact of these potentially coupled

biotic-abiotic geochemical reactions and their response to dynamic hydrologic conditions, further geochemical analyses (*e.g.*, confirmation of S intermediates and Fe(III) where sulfate reduction was active) and field examination were needed.

Building on this previous research[52,53] the present study provides a comprehensive suite of observations that demonstrates geochemical responses to changes in hydrologic gradients through time and space, including the occurrence of cryptic S biogeochemical cycling. The previous year's modeling study focused on conditions that were dominated by upwelling. For this study, we aimed to quantitatively and qualitatively characterize sediment and porewater geochemistry before, during, and after fluctuations in flux direction, with the expectation that downwelling conditions would supply more energetically favorable electron acceptors (*e.g.* oxygen, nitrate, Fe(III) oxides) and high sulfate concentrations (due to industrial pollution) from the stream water. We also sought to identify if conditions in the stream channel compared to the organic-rich, flanking riparian wetlands resulted in different geochemical responses to these fluctuating flux directions and sulfate-loading. To do this, hydrologic gradient as well as sediment, surface water, and porewater geochemistry was characterized at multiple points throughout the 2017 spring-summer-fall season at a sulfate-impacted stream flanked by wetlands in northern Minnesota (USA). X-ray absorption spectroscopy (XAS) was used to characterize the iron and sulfur speciation of sediments collected throughout the season to provide evidence for cryptic sulfur cycling, such as intermediate-valence sulfur compounds and Fe(III) phases co-existing in anoxic conditions. Our results provide a greater understanding of how the hyporheic zone geochemically responds to dynamic hydrologic conditions and the potential impact on water quality and atmospheric methane fluxes.

Materials and methods

Field site overview

Sediment cores, surface waters, porewaters, and hydrological data were collected from Second Creek, a SO_4^{2-}-rich stream flanked by wetlands near Hoyt Lakes, Minnesota (USA), as described in Ng *et al.*,[52,53] (Fig. 1). The high SO_4^{2-} concentrations (reaching as high as 10 mM) in the surface waters are largely due to the discharge of waters from upstream iron mining rock waste pits.[54] These rock-filled pits contain minerals associated with taconite mining in northern Minnesota (including hematite, siderite; goethite, and Fe sulfides),[55] and weathering of this waste-rock produces waters with elevated concentrations of SO_4^{2-} that influence the biogeochemistry of Second Creek. During the 2017 field season, relatively sparse macrophytes including wild rice (*Zizania palustris*) were growing in the sand-dominated stream channel where our field work was conducted. The riparian wetlands were much more densely vegetated with grasses and sedges, and sediments are rich in organic matter along with finer sediment particles.

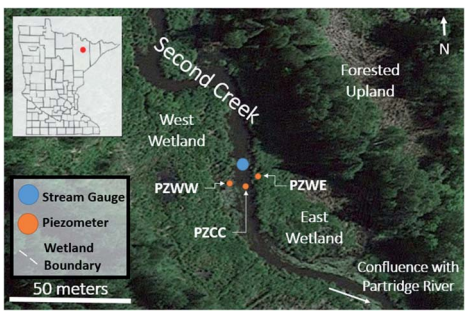

Fig. 1 Second Creek study area showing the stream channel and riparian wetlands in northeastern Minnesota (inset) with the specific locations of groundwater piezometers (orange circles), stream gauge (blue circle) and wetland boundaries (white dashed lines). Samples for sediment and aqueous geochemistry were taken adjacent to the piezometers for the west wetland (PZWW), east wetland (PZEW), and stream channel center (PZCC). This figure is based on Fig. 1 in Ng *et al.* (2017).

Hydrologic monitoring

Piezometers and stream gauges were installed from June 9 through October 21, 2017 to measure shallow groundwater head and surface water head, respectively. Three piezometers were manually installed at approximately 1 m depth: one in the center of the stream channel, one in the east wetland, and one in the west wetland (Fig. 1). The piezometers and the stream gauge contained Schlumberger Baro or Diver pressure transducers. The raw pressure data were compensated using air pressure records from Eveleth, MN (U.S. Local Climatological Data (LCD) from the National Centers for Environmental Information, https://www.ncei.noaa.gov/) and then calibrated to manual water level measurements and surveyed top-of-casing elevations taken at the start and end of the monitoring period. Resulting daily shallow groundwater and surface water head levels were used to calculate vertical hydraulic gradients and infer flux directions and magnitudes.

Sample collection overview

Samples for aqueous geochemistry of porewaters and surface waters as well as solid-phase geochemistry of sediments were collected in a transect across the 2–3 m stream, including the east and west riparian wetlands flanking the main channel, in campaigns beginning May 24, June 21, August 1, August 29, and October 20, 2017. For the campaign on August 1, most of the samples collected are assumed to reflect sediment conditions during the end of July, and therefore this sampling period is hereafter referred to as 'July'. Geochemistry samples were collected near the groundwater piezometers to understand the relationship of hydrologic flux to the hyporheic zone biogeochemistry. Sediment cores approximately 25–30 cm in length were retrieved using an HTH sediment gravity corer (Pylonex, Sweden) with 7 cm diameter polycarbonate tubes as described previously.[52] Separate cores were retrieved from each location for porewater and sediment geochemistry analysis.

Sediment sample collection. Each core for sediment sample collection was frozen with dry ice, transferred to a portable glove bag purged with nitrogen gas (N_2), and subsampled in the field. Cores were subsampled at three depth intervals, at approximately 4 cm, 10 cm, and 20 cm below the water–sediment

interface. Individual sediment subsamples were collected for analysis of acid-volatile sulfides (AVS) and solid-phase Fe and S speciation. For preservation of AVS, 5 mL of 1% (w/w) $ZnCl_2$ was added to VOA glass vials with 0.125 in septa before adding approximately 5–10 g sediment. Subsamples for Fe and S speciation were collected into vials without any preservative. After adding sediment and homogenizing by hand, all vials were purged with N_2, sealed in mylar bags with an AnaeroPack® sachet, and stored on dry ice during transport to the lab. In the lab, samples were stored at −80 °C prior to analysis.

Porewater collection. Sediment porewaters were collected using two distinct methods: multi-chambered equilibrium dialysis samplers ("peepers") that allow for high vertical resolution but 2–3 weeks averaged temporal resolution, and Rhizon samplers that enable instantaneous temporal resolution but have lower spatial resolution. The peepers collected filtered porewaters (0.2 μm filter membrane) and surface waters at or just above the sediment–water interface (when water levels were high enough) in wells spaced at 1.56 cm vertical intervals down to a depth of approximately 30–40 cm below the sediment–water interface. Peepers were filled with N_2-purged deionized water for several days before field deployment and were inserted vertically into the stream channel and wetlands (two peepers at each site, with some exceptions of when peepers failed) for two to three weeks to equilibrate prior to sampling following the general protocols by Koretsky *et al.*[56] To collect porewater geochemistry that reflect *in situ* conditions at the time of sediment core sampling (*i.e.*, instantaneous temporal resolution), 10 cm Rhizon filters with pore size ∼0.15 μm (Rhizosphere Research Products, Wageningen, NL) were inserted vertically into the sediment cores designated for porewater collection. Rhizon-sampled porewaters were collected from a sediment core top interval (∼0–10 cm) and sediment core bottom interval (∼13–23 cm), and thus represent the averaged concentration of analytes within those intervals. The outlet of the Rhizon samplers were attached with a needle to vacuum-evacuated glass serum vials sealed with butyl rubber stoppers to draw porewater from the core. Plastic wrap was secured to the top of the core to prevent oxygen diffusion into the sediment during sampling. Porewaters collected by peepers and Rhizons were analyzed for dissolved sulfide (preserved with 1 mL of 1% (w/w) $ZnCl_2$), Fe (preserved with 1–2 drops 6 N HCl), dissolved cations and anions, and alkalinity (peepers only). Additional samples were collected from peepers for methane analysis (preserved with Na_3PO_4) and pH measurements. All porewater samples were immediately stored at 4 °C until analysis.

Geochemistry analysis

Porewater and surface water chemistry. All subsurface porewaters (and surface waters near the sediment–water interface collected from peepers) were analyzed for pH, alkalinity, and dissolved anions and cations. Peeper-collected porewater samples included additional measurements of dissolved sulfide, Fe, and methane that were not measured in Rhizon-collected waters. Porewater pH was measured in the field using a pH electrode immediately following collection. Alkalinity was measured within 2–3 days upon sample collection using colorimetric acid titration. Anions (Cl^-, NO_3^-, NO_2^-, Br^-, SO_4^{2-}) were measured *via* ion chromatography (IC) using a Metrohm Professional IC Vario 1 with 4.25 mmol L^{-1} Na_2CO_3 eluent. Because the IC analysis occurs in oxic conditions, the high concentration of dissolved Fe in Second Creek porewaters (up to 85 mg L^{-1}; Ng *et al.*[52,53]) raised the concern of Fe oxide precipitation, which can adsorb phosphate. To avoid this, phosphate was measured separately on the UV-Vis spectrophotometer using an anoxic colorimetric method adapted from Hansen and Koroleff.[57] Dissolved sulfide was measured spectrophotometrically at 670 nm after reaction with a mixed diamine reagent.[58] Dissolved cations (Na^+, K^+, Mg^{2+}, Ca^{2+}) were measured *via* IC using the Metrohm protocol AN-C-103. Dissolved Fe(II) and total Fe (Fe_T) were measured spectrophotometrically at 562 nm (Agilent Cary 60 UV-Vis) using a modified colorimetric ferrozine method,[59,60] where Fe_T was measured by reducing Fe(III) to Fe(II) using hydroxylamine hydrochloride. Dissolved Fe(III) was calculated as the difference between Fe_T and dissolved Fe(II). Methane was measured on peeper samples preserved with $ZnCl_2$ by gas chromatography following the method outlined in Ng *et al.*[52]

Unfortunately, geochemical analysis showed that the "surface water" peeper samples from the sediment–water interface were likely impacted by anoxic groundwater intrusion into the peeper wells (potentially due to the hydrologic conditions or from changes in sediment suspension and vertical movement of the sediment–water interface after peeper emplacement). The actual surface water chemistry was assumed to be similar to that determined for samples previously collected using Nalgene bottles at this same field site[52] and during June of the present study.

Acid volatile sulfides (AVS). AVS was measured using a method modified from Allen *et al.*[61] that assesses the fraction of sediment sulfide available for (bio)geochemical reactions. For AVS extractions, ∼1 g $ZnCl_2$-preserved sediment was added to anoxic 6 N HCl under N_2 flow. The gas was passed through a $AgNO_3$ trapping solution for ∼45 minutes. The resulting precipitate was filtered, dried at 100 °C, and total sediment AVS was determined gravimetrically.

X-ray absorption spectroscopy (XAS)

All synchrotron-based XAS data collection was performed at the Advanced Photon Source, Argonne National Laboratory (Lemont, IL). Sediment samples were kept anoxic and shipped frozen to the facility inside N_2-purged, sealed mylar bags and were thawed immediately before analysis. At the beamline, samples were prepared inside an anoxic chamber and transported to the beamline inside N_2-purged mason jars to limit sample oxidation. Fe K-edge XAS spectra (X-ray absorption near edge structure; XANES, and extended X-ray absorption fine structure; EXAFS) and S K-edge XANES spectra were acquired from hydrated sediments collected at multiple depths and analyzed as described in the ESI.‡ Linear combination fitting (LCF) of the individual normalized spectra was conducted using a library of Fe and S reference compounds to identify and quantify structural components in sediment samples as further described in the ESI.‡

Results

Hyporheic flux

The Second Creek stream and wetlands (including field sampling subsites) had inundated conditions throughout the summer of 2017. Vertical hyporheic flux directions were determined based on hydraulic head gradient signs calculated from surface water and subsurface water head measurements, and stronger fluxes were inferred for times with higher magnitude gradients. Hydraulic head gradients in Fig. 2 showed that the vertical flux between the surface water and groundwater was dynamic (in terms of direction and

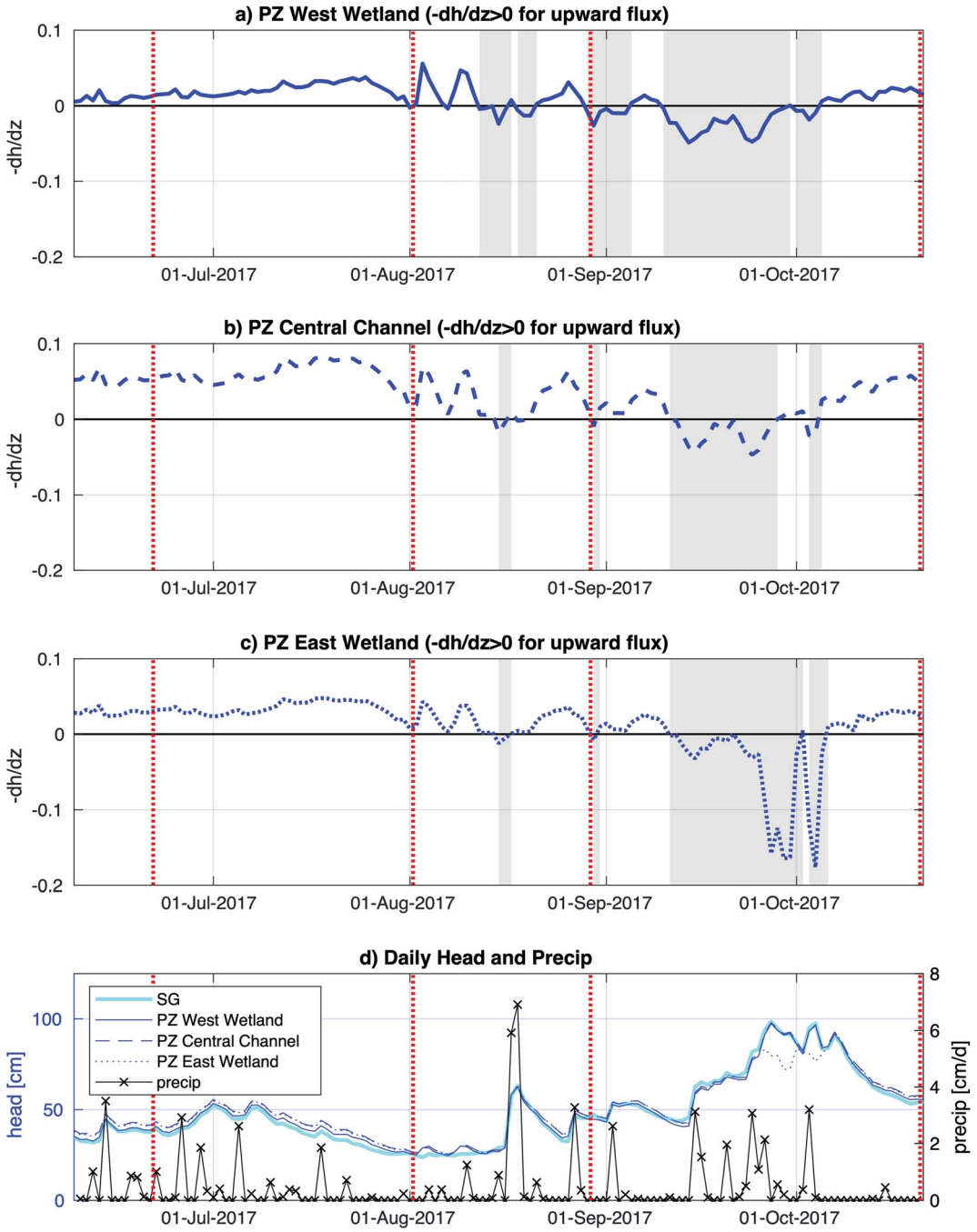

Fig. 2 Vertical hydraulic head gradient for (a) the west wetland, (b) stream channel (center), and (c) east wetland, where the white (gray) shading demarcates times with upward (downward) flux. (d) Daily precipitation from Embarrass, MN (downloaded from the Midwestern Regional Climate Center, https://mrcc.purdue.edu/) and head levels, relative to the sediment–water interface in the channel, for surface water at the stream gauge ("SG"), the west wetland shallow groundwater ("PZ West Wetland"), the central channel shallow groundwater ("PZ Central Channel"), and the east wetland shallow groundwater ("PZ East Wetland"). Dashed red vertical lines indicate sampling time points for sediment and aqueous geochemistry analyses.

magnitude) with similar seasonal changes throughout all three subsites. Upward flux prevailed at all locations from June until mid-August, when a strong precipitation event caused a brief flux reversal. From mid-August to early October, flux directions and magnitudes were variable, with rainier conditions triggering multi-day to multi-week downward flux periods. Steady upward flux conditions then resumed for the rest of October. The central channel and west wetland hydraulic gradient time series show quite similar trends over time and magnitudes in 2017, consistent with similarities found between the two locations by Ng *et al.*[52,53] in previous years, but the east wetland, which had not previously been monitored, showed some distinct differences. In particular, the east wetland showed two multi-day periods in late September to early October with very negative hydraulic gradient that suggest strong downward fluxes.

Surface and porewater geochemistry

pH and nutrient trends. The pH of the peeper-collected surface and pore waters was near neutral at all three sampling locations (west and east wetlands and stream channel) throughout the summer 2017 season (Table S1 and Fig. S1‡). There were no substantial pH differences among the sampling locations or months, but the surface water pH (~7.5–8.0) was generally higher than the porewater pH (~6.9–7.3). Porewater pH decreased as sediment depth increased within the first 10 cm below the sediment water interface, but then stayed generally consistent deeper in the sediment profile at all locations.

Compared to peeper-collected surface water phosphate (PO_4^{2-}) (mean of 0.07 ppm), subsurface porewater PO_4^{2-} levels were higher at all locations throughout the summer, with concentrations substantially higher (by 1–2 orders of magnitude in peeper samples) in the riparian wetlands than in the stream channel (Table S1 and Fig. S2‡). In contrast to PO_4^{2-}, concentrations of NO_3^- + NO_2^- in the surface water and subsurface porewaters at all three locations were all consistently very low throughout the summer, with levels that were often too low to be accurately measured (Table S1 and Fig. S3‡). These trends are described in greater detail in the Results section (nutrient trends) of the ESI.‡

Methane trends. Dissolved CH_4 was detected at all Second Creek subsites in the sediment porewaters collected from peepers. There was variability between the replicate peepers for each month due to spatial heterogeneity in the subsurface, but general trends show porewater concentrations consistently higher in the riparian wetlands than in the stream channel, particularly during the warmer months of June, July, and August (Table S1‡ and Fig. 3). At all locations and sediment depths, porewater CH_4 concentrations were generally lowest in October. Methane concentrations in the west wetland were highest in the upper 10 centimeters in June, July, and one August peeper, but peaked at greater depths (10–15 cm) in August at the other peeper location. Average methane concentrations across all depths in the west wetland were 0.29, 0.28, 0.23, and 0.08 mM, for June, July, August, and October, respectively. As opposed to the west wetland, vertical trends in porewater CH_4 were less consistent in the east wetland. Concentrations in July were generally highest (ranging from 0.06 to 1.5 mM, mean = 0.51 mM across all depths and peepers) and decreased only minimally at depths ~30 cm, whereas CH_4 levels peaked in August (~0.4 mM) at about 12–18 cm depth with a concentration profile similar to that in the west wetland and an average methane concentration of 0.20 mM across all depths. The lower CH_4 concentrations in the stream channel (~0.5 mM CH_4) generally peaked between 5 and 15 cm sediment depth during June and July, and decreased to an average of ~0.016 mM by about 30 cm depth in all sampling months. Average methane concentrations in the channel across all depths were 0.10, 0.04, 0.06, and 0.14 mM for June, July, August, and October, respectively. Although CH_4 levels were generally low (or not measured) in surface waters during most months, samples collected in August appeared elevated in at least one peeper per location

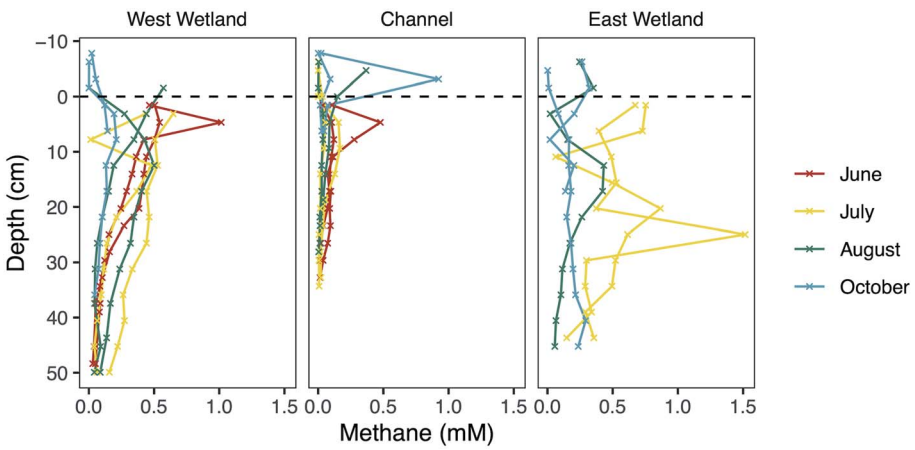

Fig. 3 Methane concentrations (mM) of surface waters and porewaters with depth (cm) at Second Creek in the stream channel and both wetlands collected from replicate peepers during June, July, August, and October, 2017. Zero depth (indicated by the black dashed line) marks the sediment–water interface, increasingly positive depth values indicate deeper sediment depth, and negative values indicate surface water. Peepers were not installed in May or in the east wetland in June.

with concentrations above the sediment water interface as high as 0.35 mM in the east wetland, 0.37 mM in the channel, and 0.57 mM in the channel. October also saw elevated surface water CH_4 concentrations in one peeper from the east wetland (0.26–0.32 mM) and one possibly spurious measurement (0.92 mM CH_4) in the channel.

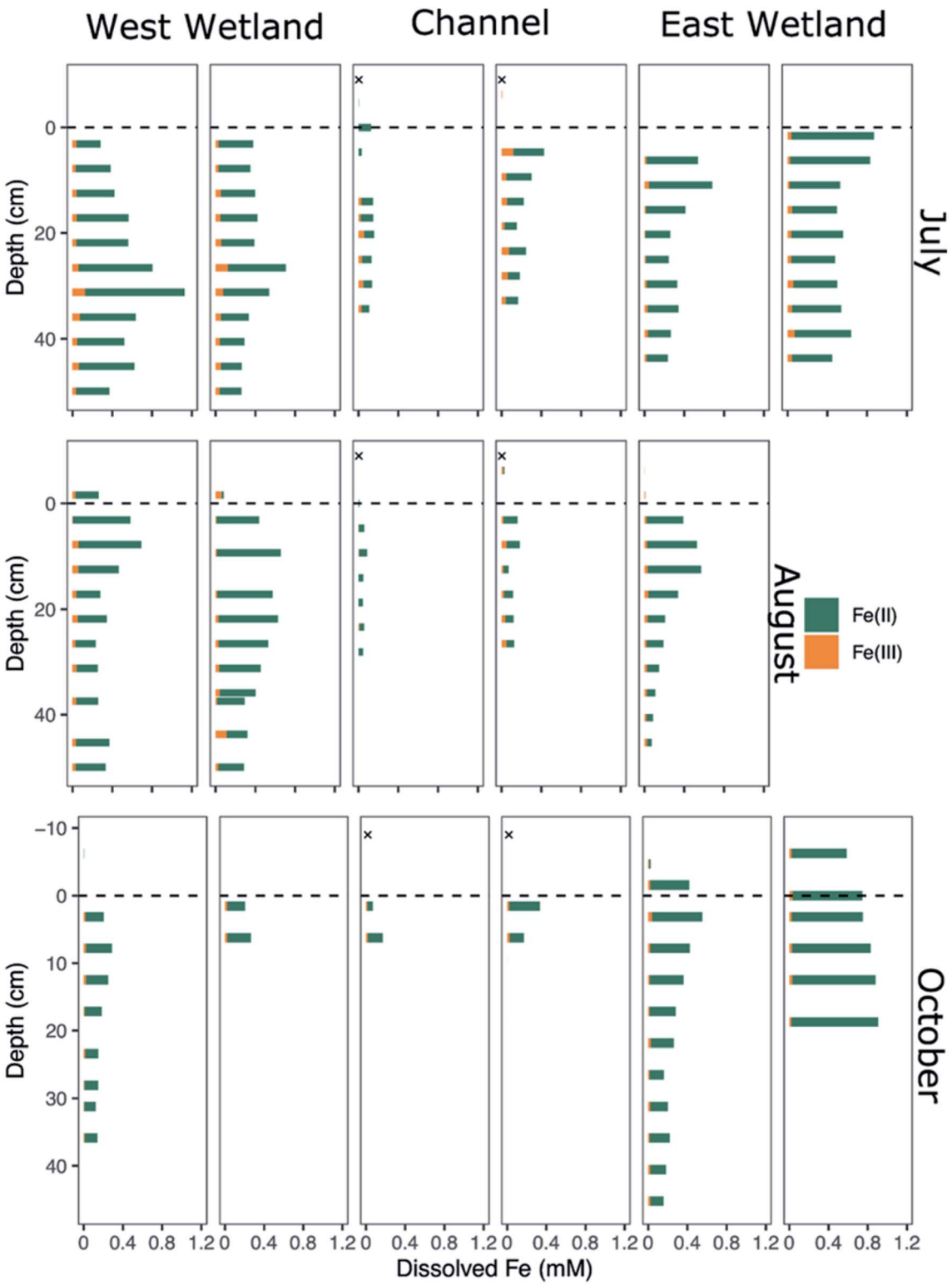

Fig. 4 Dissolved iron speciation and concentration of porewaters collected from replicated peeper samplers as a function of sediment depth for July (top), August (middle), and October (bottom). Measured Fe(II) concentrations are denoted by green bars and calculated Fe(III) concentrations are denoted by orange bars. Channel surface water samples indicated by an "X" were collected in 2015 *via* Nalgene bottles[46] and represent nearer surface conditions compared to peeper samples taken just above the sediment water-interface.

Dissolved iron concentration and speciation. At all locations, Second Creek surface water typically contained low (<20 μM) dissolved total Fe concentrations compared to porewaters[52] (Fig. 4). "Dissolved" Fe could include aqueous Fe compounds, organically complexed Fe, or colloidal or amorphous Fe (small enough to pass through a 0.2 μm filter). While porewater total dissolved Fe varied substantially based on depth, location, and month sampled, the highest concentrations were measured in the wetlands. In the channel, porewater Fe concentrations typically ranged between 50-150 μM, though a few measurements were substantially higher (~400 μM). In the east and west wetlands, porewater Fe concentrations generally peaked at depths between 5–15 cm with concentrations ranging between 570–870 μM. One notable exception was the west wetland in July, where the highest observed total Fe concentrations (707 μM and 1128 μM) occurred lower in the sediment profile (26–31 cm). Overall, dissolved Fe in the west wetland and channel porewaters decreased significantly from August to October, while the east wetland porewater total Fe remained relatively similar during all three months. Interestingly, there are high concentrations of Fe in the surface waters measured from some of the peepers (*e.g.*, most of the October timepoints), which potentially indicates suspension of small/colloidal particles (small enough to pass through a 0.2 μm filter), upwelling of anoxic water from depth, or vertical movement of the sediment water interface throughout the 3 weeks sample collection period.

Though Fe(II) was the dominant dissolved Fe species in all water samples including surface waters, dissolved Fe(III) was present in nearly all porewater samples, generally comprising between 10-20% of the total dissolved Fe, with several samples containing up to 25–40% (Fig. 4). In general, dissolved Fe(III) was a greater relative proportion of the total dissolved Fe concentrations in the channel porewaters than in either of the wetlands.

Aqueous sulfate trends. Aqueous SO_4^{2-} concentrations were elevated in Second Creek surface waters relative to porewaters as well as to regional non-impacted surface waters[62] during all months sampled, particularly in the stream channel. Surface water levels measured as high as ~5 mM SO_4^{2-} in July and August 2017 in the stream channel when collected from peepers (Fig. 5 and Table S1‡). This is within the range of upper surface

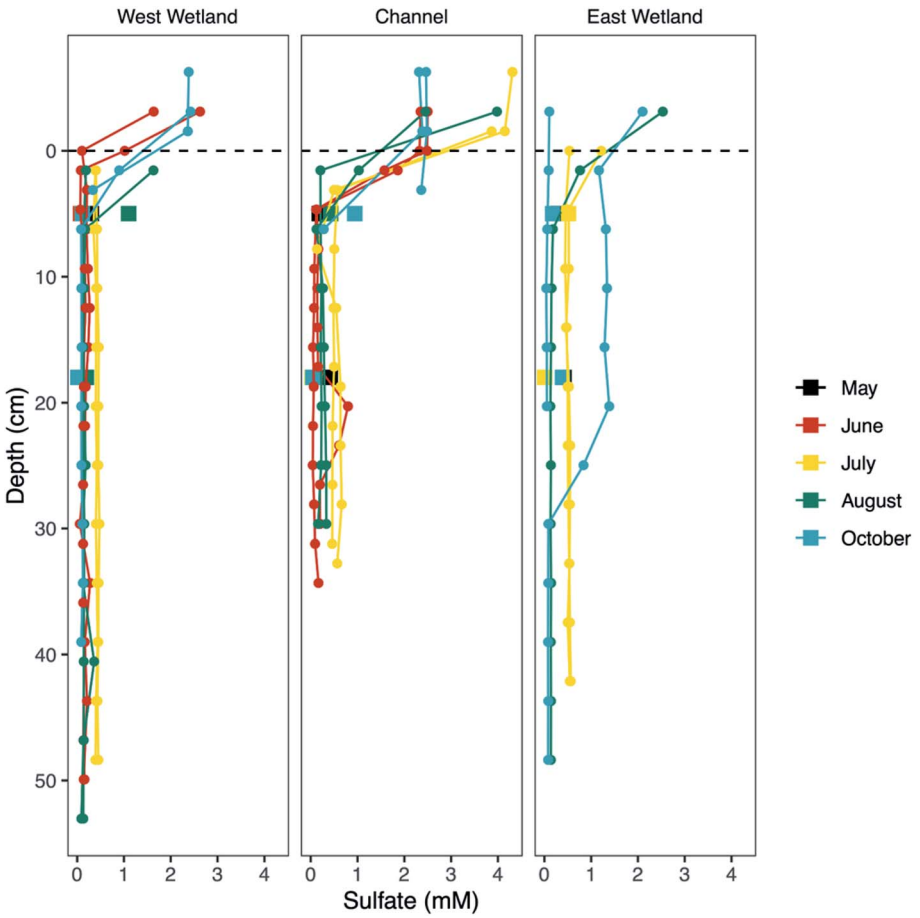

Fig. 5 Dissolved sulfate concentration *versus* sediment depth at Second Creek in the stream channel and both wetlands during May, June, July, August, and October of 2017. Small solid circles represent 'peeper' porewater data whereas porewaters collected using Rhizon samplers are denoted by filled squares and represent averaged concentrations from 0–10 cm and 10–20 cm sediment depth. Zero depth (indicated by the black dashed line) marks the sediment–water interface, increasingly positive depth values indicate deeper sediment depth, and negative values indicate surface water. Peepers were not installed in May or in the east wetland in June.

water values measured in the two wetlands and channel during June 2017 (2.9–7.8 mM) and in previous summers (2.8–10 mM).[52] In the wetlands, there was insufficient ponded surface water midsummer for sampling, but at other times, wetland surface water SO_4^{2-} concentrations from peepers averaged 2.32 mM SO_4^{2-} (west wetland) and 1.58 mM SO_4^{2-} (east wetland).

Concentrations of porewater SO_4^{2-} rapidly decreased with increasing depth up to ~5 cm in the channel and west wetland to 0.1–0.5 mM and 0.06–0.4 mM, respectively (Fig. 5). Below 5 cm, average SO_4^{2-} concentrations were consistently low during the months of June and August in the channel (0.14 mM) and west wetland (0.21 mM), and slightly higher (0.50 mM and 0.43 mM, respectively) in July. East wetland porewater SO_4^{2-} showed similar trends in July and August where concentrations decreased over the top ~5 cm to 0.52 mM in July and 0.18 mM in August. In October, the replicate peepers in the east wetland indicated differing trends of porewater SO_4^{2-}; one had low concentrations (≤0.05 mM) throughout the sediment profile, while the other showed concentrations in excess of 1.2 mM before decreasing to ~0.08 mM below ~30 cm. Rhizon samples generally showed slightly higher concentrations of SO_4^{2-} relative to the peeper samples for the upper 10 cm, but concentrations at greater depths were consistent between the two sampling methods.

Porewater sulfide trends. Concentrations of porewater aqueous hydrogen sulfide ($H_2S + HS^- + S_2^-$) were low in all three locations, with an overall average porewater concentration of 3.2 μM. Average sulfide across all depths during the different sampling months ranged from 3.2–4.8 μM in the west wetland, 2.2–4.9 μM in the channel, and 2.1–5.5 μM in the east wetland (Fig. 6). In June and August, sulfide concentrations remained fairly constant throughout the sediment profile, while there were higher concentrations and greater spatial variability in July and October. In the west wetland, for example, sulfide concentrations spiked to 22 μM around ~31 cm depth during July for one of the peeper replicates, while the other July peeper in the west wetland remained between 0.25–2.5 μM at all depths measured. In the channel, the highest concentrations were measured in July at or just below the sediment water interface, reaching as high as 22.5 μM in one peeper at 0 cm depth and 11 μM at 4.7 cm in the other peeper. In the east wetland, aqueous

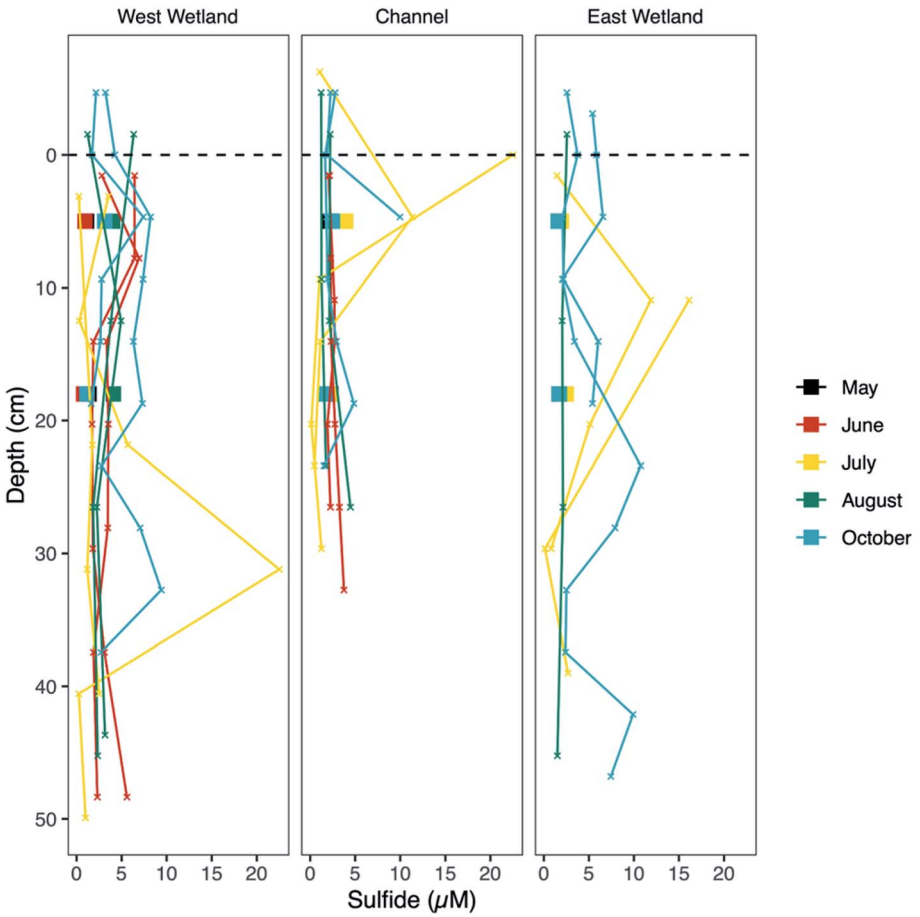

Fig. 6 Dissolved sulfide concentration *versus* sediment depth at Second Creek in the stream channel and both wetlands during May, June, July, August, and October of 2017. Small "x's represent 'peeper' porewater data whereas porewaters collected using Rhizon samplers are denoted by filled squares and represent averaged concentrations from 0–10 cm and 10–20 cm sediment depth. Zero depth (indicated by the black dashed line) marks the sediment–water interface, increasingly positive depth values indicate deeper sediment depth, and negative values indicate surface water. Peepers were not installed in May or in the east wetland in June.

sulfide concentrations were at a maximum (12 and 16 μM) in both peepers at ~10 cm depth, with lower concentrations deeper in the sediment profile. Vertically averaged sulfide concentrations collected from Rhizon samplers were generally lower than peeper porewater concentrations across all sites at both depth intervals and did not show substantial variability with depth or sampling period.

Sediment AVS fraction

Measurements of unfiltered acid volatile sulfide, possible phases include reactive FeS and ZnS phases and free sulfide, represent the fraction of sulfide in the sediments readily influenced by biogeochemical cycling and fluctuations in redox gradients in the hyporheic zone. The average unfiltered AVS fraction (Fig. 7) was comparable in all three locations (30.0, 25.5, and 25.6 μmol S g^{-1} sediment, for west wetland, channel, and east wetland, respectively). There was a general decrease in AVS with increasing sediment depth with some exceptions (*e.g.*, AVS in the west wetland was highest in the bottom interval in July and October).

Molecular-scale characterization of sediment iron

Fe K-edge X-ray absorption spectroscopy (XAS) was performed to characterize solid-phase Fe at different depths in the sediment profile. Linear combination fits of the Fe K-edge XANES spectra showed that all samples were mixed valence, with average Fe oxidation states ranging between 2.14 and 2.67 (Table S2 and Fig. S4‡). To determine individual contributions of Fe species, Fe K-edge EXAFS were also analyzed *via* LCF. Based on principal component analysis (PCA), most of the spectral variance (93.4–94.8%) in the 42-spectra dataset was explained by 4–5 components. Target transform analysis found that all reference compounds were potentially suitable components; therefore, all reference standards (Fig. S5‡) were used in the LCF fitting process. Optimal fits (Fig. S6‡) were selected using the minimum number of components, with additional components added only if the additional components improved the *R*-factor by 15% or more.

Although there was substantial variability in the Fe-containing solid phases collected from the three locations across space, depth, and time, some trends were apparent (Fig. 8). Nearly all samples contained phyllosilicate-bound Fe (8–52%), as well as in sulfide phases including iron monosulfide (0–45%), iron disulfide (0–23%), and/or pyrrhotite (0–22%). Fe(III)-containing oxides such as hematite, goethite, and ferrihydrite were identified (between 0% and 31% maximum) in some locations and sampling times; no Fe(III)-containing oxides were detected in July. Dissolved or adsorbed Fe(II) and Fe(III) were also present in all samples (combined relative abundance 16–62%), with dissolved or adsorbed Fe(III) present at all depth intervals for most samples collected. The speciation of the dissolved and adsorbed Fe(II) and Fe(III) components, the Fe(III) phyllosilicate, and iron disulfide is described in greater detail in the ESI.‡

One primary difference between the channel and wetlands was that while the vast majority of sulfide-containing Fe in the channel was identified as iron monosulfide and pyrrhotite, sulfide-containing Fe in the wetlands typically contained a mixture of iron monosulfide, pyrrhotite, and iron disulfide, especially deeper in the sediment profiles. As opposed to the other iron sulfides, pyrrhotite in this system is likely detrital due to its relatively common presence in geologic formations in the region.[63] The range of total sulfide-containing Fe relative abundances in the three locations, however, were quite similar (west wetland = 16–47%, channel = 20–49%, east wetland = 16–53%). While iron disulfides in the wetlands increased with depth, the total relative abundance of iron sulfide phases in the wetlands tended to decrease or stay the same with increasing depth at most sampling times except June (west wetland). Conversely, the channel total sulfide abundance showed more variability but with no clear patterns across sampling times.

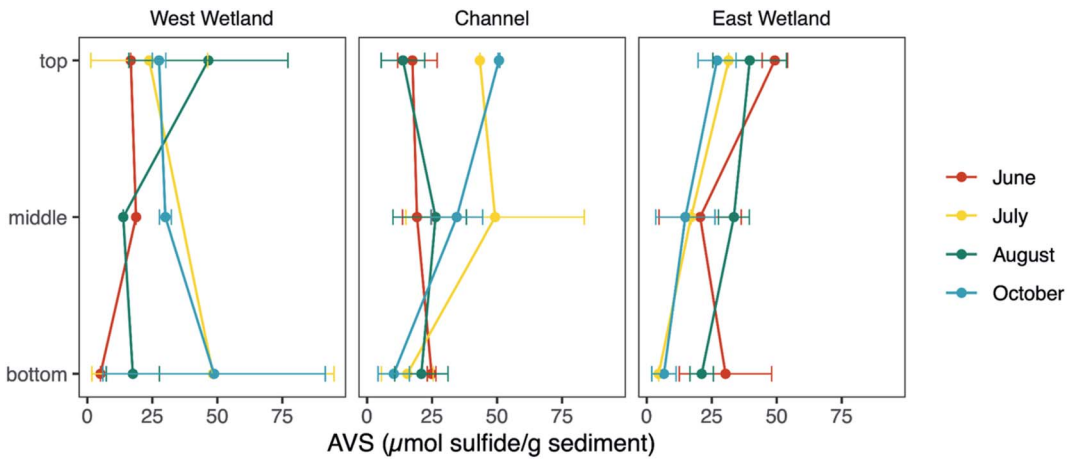

Fig. 7 Measurement of unfiltered acid volatile sulfide (AVS) fraction in hyporheic zone sediments collected from the west wetland, channel, and east wetland (left to right, respectively) of Second Creek during the 2017 field season. The average AVS fraction is denoted by filled circles. Error bars represent the range of values measured for replicate samples. The 'top', 'middle', and 'bottom' correspond to sediment core intervals centered at approximately 4 cm, 10 cm, and 20 cm depths, respectively.

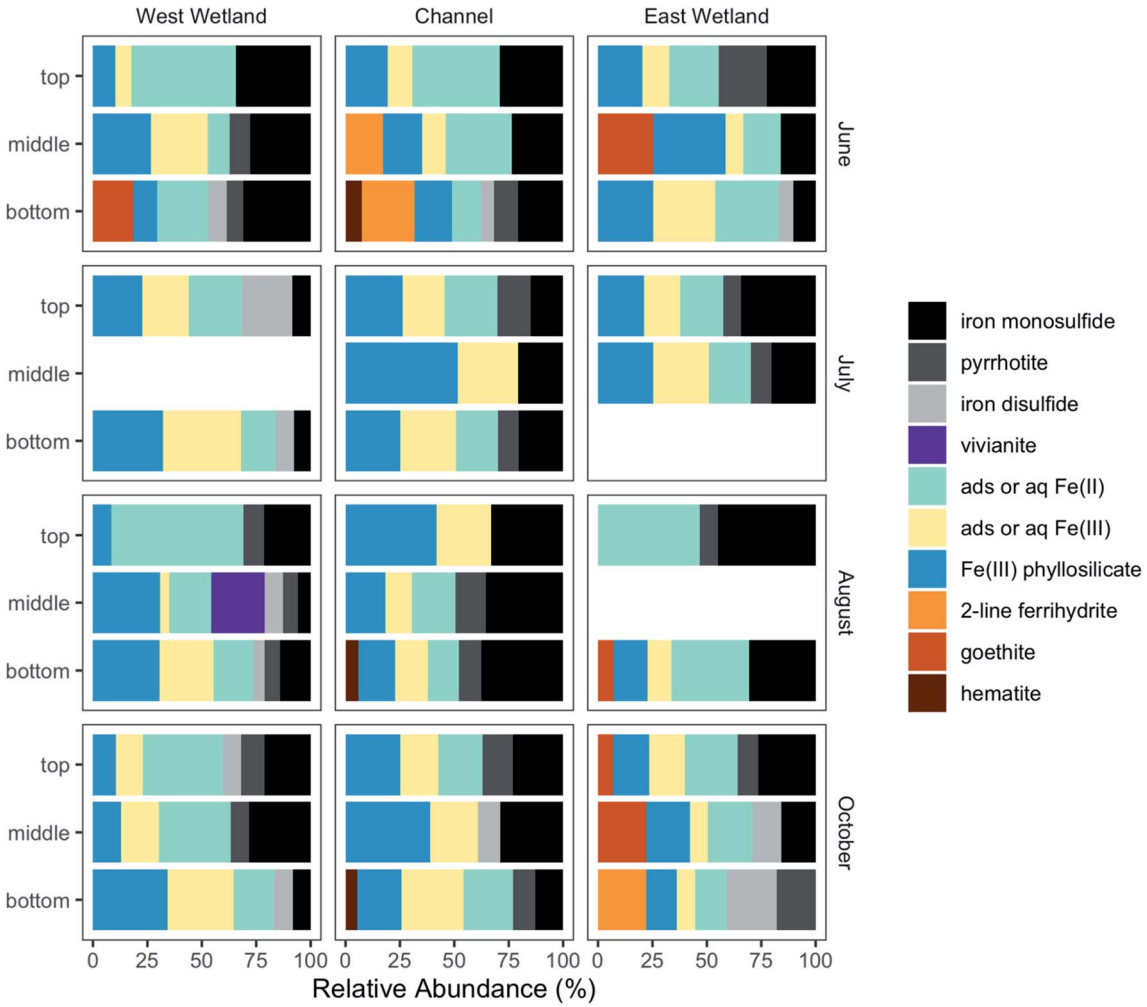

Fig. 8 Linear Combination Fits (LCF)s of Fe K-edge EXAFS spectra from sediments collected from the west wetland, stream channel, and east wetland of Second Creek showing relative abundance for Fe-containing components in bulk sediment samples with depth through the 2017 sampling season. For LCFs, the sum of all Fe species was normalized to 100%. The 'top', 'middle', and 'bottom' correspond to sediment core intervals centered at approximately 4 cm, 10 cm, and 20 cm depths, respectively.

Molecular-scale analysis of sediment sulfur

Characterization of S-bearing solid phases was performed using S K-edge XANES. Principal component analysis indicated that 99.4% of the variance of the 22 S-XANES sample spectra could be fit using 4 components; adding the fifth component increased the cumulative variance to 99.8%. As with the Fe-EXAFS, target transform analysis indicated that all reference standards (Fig. S7‡) were potentially acceptable components; therefore, all reference standards were used in all LCF fits. Optimal fits were selected using the minimum number of components with additional components included only if they improved the *R*-factor by at least 15%.

Nearly all sample spectra contained peaks in two primary regions (Fig. S8‡). Peaks in the 2470–2475 eV region result from more reduced S species including organic sulfide (*e.g.*, methionine, cystine), inorganic sulfide (*e.g.* mackinawite (FeS), chalcopyrite), and elemental sulfur, while peaks in the 2480–2485 eV region result from more oxidized S components (*e.g.* sulfate, thiosulfate). S-XANES spectra were fit using reference standards selected to represent a range of S oxidation states ranging from −2 to +6, including a mix of inorganic and organic compounds (Table S3 and Fig. S7‡).

S XANES detected a wide range of S species and valence states in the Second Creek sediments, and nearly all sediment samples contained a mixture of S valence states. In the west wetland and channel, oxidized S forms (inorganic sulfate and organic oxidized S; oxidation indices = +5 − +6) were relatively consistent with depth, between ~6–20% of the total S pool. The east wetland was more variable with respect to oxidized S content. For example in June, there was 0% oxidized S detected in the top interval and 65% oxidized S in the bottom interval. Both June and August saw increased oxidized S with increasing depth in the east wetland, while the greatest relative abundance of oxidized S in October occurred in the middle depth. In all locations, intermediate oxidation state S forms (organic with oxidation states +1 − +2, inorganic with mixed oxidation states of −1/+5 and −0.1/+3.3, S(0), and polysulfide with an oxidation

state of −1.3/−0.1) were detected at nearly all depths, with relative abundances up to 46% organic forms and 28% inorganic mixed valence compounds. Note that while polysulfide also has a mixed valence state, it's chemistry in the hyporheic zone is expected to be different than the other inorganic mixed valence compounds (thiosulfate and tetrathionate) and was therefore not included in the inorganic mixed valence group. Elemental S (oxidation state = 0) and polysulfide (oxidation state = −1.3/−0.1) were also significant components of many of the sediment samples in all months and all locations, especially August (all locations, 35–66%), June in the west wetland (43–56%), and October in the east wetland (30–63%). Regarding the most reduced S phases, inorganic metal sulfides were also identified in most samples. In general, the metal sulfides were identified as iron monosulfide (mackinawite, oxidation state = −2), though iron disulfide (*e.g.*, pyrite, oxidation state = −1) was identified in the bottom samples of the east and west wetlands (54 and 37%, respectively) in October and in the channel (31%) during August. Contrary to iron disulfide, iron monosulfide was typically most abundant in the shallower depths of the cores and the relative abundance decreased with increasing depth.

Discussion

Impact of hydrology on hyporheic zone geochemistry

A major goal of this study is to examine how major changes in vertical hyporheic flux impacts the subsurface geochemistry of riparian wetland and stream environments through time. Though we do encompass several small shifts in hyporheic flux within our sampling period, we notably do not observe major changes in porewater geochemistry after shifts from upward to downward fluxes, except for increased sulfate and Fe oxides in the east wetland in October (Fig. 5) at Second Creek. This suggests that the system can recover quickly to previous conditions and/or the system is buffered against major changes due to low magnitude or unsustained hyporheic flux changes or by geochemical controls related to metastable phase reactivity. Our data indicates that all these factors are contributing to the observed conditions. Flux magnitude variability across the wetlands and sediments is evident and can explain some of the geochemical excursions in sulfate and iron. For example, the ∼6 times greater magnitude hydraulic gradient in the east wetland in late September to early October compared to the channel and west wetland (Fig. 2) indicates that the east wetland experienced much stronger downward flux at that time which facilitated deeper penetration of elevated oxygen (resulting in Fe oxide formation) and sulfate (Fig. 5 and 8). Interestingly, previous inverse modeling at Second Creek suggests that the hydraulic conductivity of channel sediments can be up to 2–3 times greater than for wetland sediments,[53] so elevated sulfate and Fe oxides at depth was expected in the stream rather than either wetland. The strong locally measured downward flux in the east wetland could potentially result from influences of the regional west to east groundwater flow regime, which would be more likely to support losing conditions on Second Creek's east bank compared to its west bank.[52] Variable changes in geochemistry due to lateral transport has recently been demonstrated at topographic features such as stream meanders.[64]

Because the Second Creek sediments have substantial quantities of organic matter[52] that can generate highly reducing conditions, the system may be buffered against short-duration or small-scale hydrologic changes that allow only limited transport of surface water electron acceptors into the hyporheic zone relative to redox rates.[65–67] In particular, previous model simulations[52] showed that high sulfate concentrations from surface water moved only 5 cm or less into the sediment over a few days after a switch from upward to downward hyporheic flux, which is about the duration of most of the flow reversals observed in this 2017 study period. Unfortunately, the late August sampling coincided with the start of the flux reversal to downwelling conditions, which may not have allowed sufficient time to observe measurable geochemical changes, and the October sampling occurred two weeks after the strong downwelling conditions flipped to upwelling hyporheic flow. The higher porewater sulfate concentrations collected by peepers, which reflect a ∼2 weeks equilibration time, compared to Rhizons, which capture temporally instantaneous conditions, in the east wetland in October (2017) provide additional support that rates of SO_4^{2-} reduction following the resumption of upwelling conditions were rapid enough to decrease porewater SO_4^{2-} to previous values on short (<2 weeks) times scales.

Additional buffering against substantial geochemical responses during these brief hydrologic excursions could be due to the formation of redox active metastable phases (RAMPs) such as ferrihydrite and goethite, or mixed- and intermediate-valence sulfur compounds; which may only appear in "hot spots" and/or "hot moments".[45,68,69] These RAMPs have the potential to act as either/both electron donors and acceptors and their production is kinetically controlled, making their occurrence highly transient as they serve as powerful redox buffers.[45] At Second Creek, the presence of Fe and S-bearing RAMPs at various times and depths is confirmed by bulk sediment Fe EXAFS, porewater Fe speciation, and S XANES analysis. Although bulk sediment Fe EXAFS analysis indicates that solid phase Fe(III) is mostly depleted except for in June and October, the analysis revealed a substantial portion of dissolved or adsorbed Fe(III) within the sediments (Fig. 8). Further, both dissolved Fe(II) and Fe(III) are present in porewaters at all depths in the hyporheic zone (Fig. 4), as has previously been observed in both anoxic and oxic environments.[70–74] The stability of dissolved Fe, particularly Fe(III) in anoxic porewaters, is likely enhanced by complexation with various organic matter functional groups.[75–78] These organic matter functional groups may further act as electron shuttles (*i.e.*, as RAMPs), serving to oxidize or reduce Fe and other redox active analytes (*e.g.*, sulfur) through chemical or microbial processes.[45,75,79,80] The hydrologic regime changes (*e.g.*, gaining to losing stream conditions) or general mixing of oxygenated surface waters and anoxic groundwaters in the hyporheic zone at Second Creek may facilitate the rapid recycling of these inorganic and organic RAMPs,[45] which may be why metastable dissolved Fe(III) is consistently measured in the subsurface despite its likely reduction during cryptic S cycling.

Geochemical evidence for a cryptic sulfur cycle linked to Fe cycling

Cryptic S cycling has been invoked as an important process in previous modeling studies at Second Creek, though previous investigations have not shown substantiated geochemical evidence for this process. Porewater sulfate profiles measured in the current study provide one line of geochemical evidence that cryptic S cycling is occurring. Throughout the entire season, robust sulfate reduction is evident as porewater sulfate concentrations are substantially lower than surface waters (~3–10 mM sulfate). Despite an abundance of electron donor, however, the sulfate concentrations never decrease to 0 mM in the porewater (Fig. 5) and remain moderately elevated (~0.14–0.5 mM sulfate) at depth. These deep (to at least 50 cm depth) porewater sulfate concentrations are greater than what would result solely from downward dispersion and mixing of surface water (which occurs even during upwelling), as indicated by calibrated reactive transport model results for conservative solutes (see ESI‡ in Ng *et al.*[52,53]). This indicates that sulfate is being regenerated in the anoxic hyporheic zone at Second Creek, likely by sulfide oxidation processes[81–83] as part of a cryptic S cycle.

There are a variety of abiotic and biotic pathways that promote anoxic sulfide oxidation, ultimately leading to the production of sulfate that could explain this regeneration at Second Creek. In marine conditions (*i.e.*, high sulfate concentrations, ~3–10× that of Second Creek surface water), elemental S is the predominant product from sulfide oxidation coupled to Fe(III) oxide reduction, possibly formed *via* polysulfide

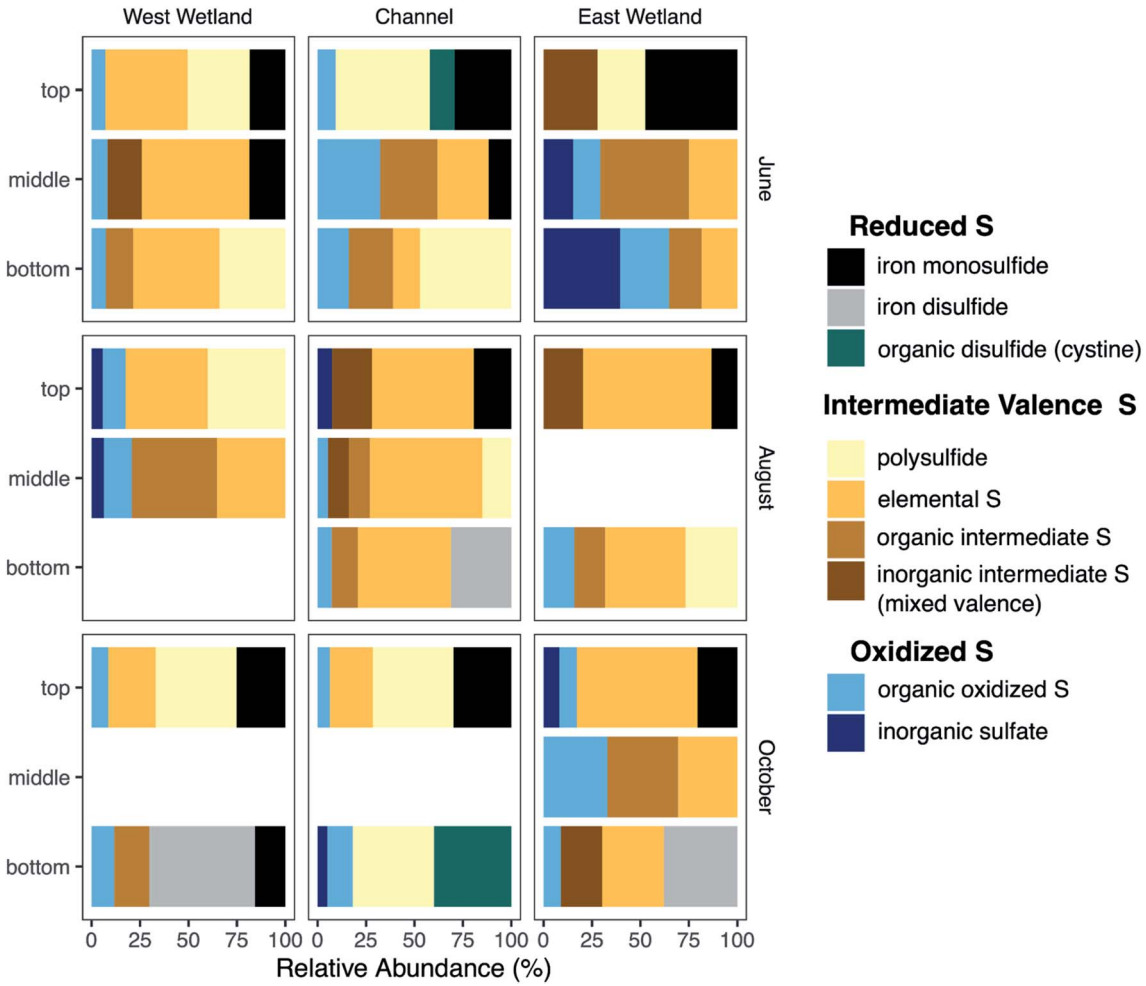

Fig. 9 Linear Combination Fits (LCF)s of S K-edge XANES spectra from sediments collected from the west wetland, stream channel, and east wetland of Second Creek showing relative abundance for S-containing components in bulk sediment samples with depth through the 2017 sampling season. For LCFs, the sum of all S species was normalized to 100%. The 'top', 'middle', and 'bottom' correspond to sediment core intervals centered at approximately 4 cm, 10 cm, and 20 cm depths, respectively. The legend is approximately ordered with more reduced S forms at the top (iron monosulfide, iron disulfide, and organic disulfide), intermediate valence species in the middle (polysulfide, elemental S, organic intermediate S, and inorganic intermediate S (mixed valence)), and more oxidized species on the bottom (organic oxidized S and inorganic sulfate). In three cases, groups of references were combined into a single category based on oxidation index. "Organic oxidized S" (oxidation index = 5–6) includes sodium dodecyl sulfate (SDS), sulfanilamide, and sulfonate references. "Organic intermediate S" (oxidation index 1–2) includes sulfoxide and thiophene reference compounds. "Inorganic intermediate S (mixed valence)" includes thiosulfate and tetrathionate (oxidation indices −1/+5 and −0.1/−3.3, respectively).

intermediates.[33,83–85] Polysulfide also forms from equilibrium with elemental sulfur and sulfide.[86] Given the high relative amount of elemental S and polysulfide in all three locations (Fig. 9) and relatively low amounts of dissolved sulfide and solid iron sulfide phases (*e.g.*, FeS or pyrite), sulfide oxidation seems to be a predominant pathway in the Second Creek hyporheic zone. Further reactions involving these S-intermediates, such as microbial S(0) and polysulfide oxidation[87,88] and disproportionation reactions[89,90] produce sulfate. High concentrations of aqueous sulfide have been shown to inhibit elemental S disproportionation,[89] but the low sulfide concentrations detected throughout the hyporheic zone porewaters in both the wetlands and stream channel suggest that disproportionation is a viable pathway for sulfate regeneration at depth.

Both solid metastable Fe(III) oxyhydroxides and organic Fe(III) complexes are likely driving sulfide oxidation, and therefore cryptic S cycling, in the Second Creek sediments. Reactive, metastable Fe(III) oxyhydroxides, such as ferrihydrite identified in our Fe EXAFS analysis, can serve as the TEA driving aqueous sulfide oxidation in anoxic environments.[33,91–93] Using an experimental laboratory approach, Taillafert *et al.*[76] further showed that Fe(III) complexed by organic matter can be reduced by sulfide at circumneutral pH. This reaction is substantially faster (half-time of seconds) than the reduction of solid Fe(III) mineral phases (minutes to hours). A follow-up study showed that Fe(III)–organic complexes present in natural sediment porewaters reacts with aqueous sulfide, potentially through a S(0) intermediate, ultimately forming FeS in low sulfide environments and pyrite in high sulfide environments.[94] Although organic–Fe(III) could not be specifically identified using the methods in this study, the measurable amounts of aqueous or adsorbed Fe(III) detected by Fe EXAFS sediment analysis (Fig. 8) and ferrozine-based valence state measurements of the porewaters (Fig. 4) suggest that Fe(III) was stabilized in the anoxic zone by organic complexes.

High quantities of DOC may also be playing a direct role in sulfide oxidation at Second Creek. Previous studies have shown that anoxic reactions between sulfide and aqueous humic acids are fast (minutes to hours) and produce S(0), thiosulfate, and a variety of different reduced organic S compounds.[95,96] While DOC speciation was not conducted as part of this study, we suspect that humic acids were present, as indicated by the yellow-brown hues in the extracted porewaters, and may be playing a role in Fe(III) complexation. Mn(III/IV) oxides can also serve as a TEA and drive cryptic S cycling,[97,98] and while these phases were not measured in the present study, elevated Mn^{2+} was detected previously in the porewaters.[52,53] Further characterization of potential electron acceptors driving sulfide oxidation at Second Creek could yield valuable information about coupled redox cycles and turnover of organic matter in these dynamic environments.

Accumulation of intermediate valence S phases in hyporheic zone sediments

Several of the reactions that drive the cryptic S cycle at Second Creek likely produce various intermediate S species. In studies that have examined S speciation, elemental S and inorganic polysulfides are detected in sediments and porewaters,[35,85,99,100] sometimes comprising a substantial portion of the total S pool depending on the environmental conditions of both terrestrial and marine ecosystems. In the present study, S XANES analysis shows that both S(0) and polysulfide are present in high relative abundance (Fig. 9), and are often the predominant S forms, throughout the ecosystem regardless of season, hydraulic flux, location, or depth (for further explanation of S XANES interpretation, see the ESI‡). The persistence of S(0) may be the result of organic matter S(0) stabilization, as has been observed recently.[101–103] A portion of the elemental S analyzed in our study could also be organic polysulfides as some organic polysulfides have absorption energies and peak features that are overlapping with, and potentially indistinguishable from, that of elemental S.[104] Organic polysulfides, though not targeted in most S speciation studies, have been found to be abundant in some anoxic environments such as salt marsh and marine sediments,[28,100,104,105] supporting the hypothesis that these compounds are present in the organic rich sediments at Second Creek, particularly in the wetlands. Identification of these compounds would require additional S XANES database refinement and alternative analytical approaches.

The relatively high contributions from intermediate valence S compounds compared to reduced iron monosulfide and disulfide forms (Fig. 9) were somewhat unexpected, given that AVS (presumably mostly FeS, as porewater sulfide measurements were low) was measured in all samples (Fig. 7), and reduced inorganic sulfides were consistently detected by Fe EXAFS analysis (Fig. 8). One possible explanation for this analytical discrepancy is that inorganic sulfides represent a relatively small (<5%) fraction of the total S pool (not quantified in this study) and were thus not detectable by S XANES. It is also plausible that the AVS results were inflated due to partial or complete extraction of non-target S phases such as organic polysulfides.[106,107]

Geochemical variability between the riparian wetland and stream hyporheic zone

Despite their close proximity, the riparian wetlands exhibit substantially different porewater geochemistry compared to the channel porewaters for some analytes, particularly P, Fe, and CH_4 (Fig. 3, 4 and S2‡). These differences likely result from higher accumulations of total organic carbon in the wetland sediments compared to the stream sediments.[52] Living and decaying aquatic plants are substantially more abundant in the wetlands than in the stream channel, although plants like wild rice (manoomin) grow sparsely throughout this section of the stream channel. This organic matter can serve as an electron donor, and even sometimes as an electron acceptor, for redox reactions that drive important biogeochemical cycles such as in N, Fe, S, and methane.[45,95,108,109] Nitrate concentrations remained low at all sampling locations in both the surface waters and porewaters, which may suggest that biogeochemical nitrogen cycling is relatively limited in this system. It is also possible that the sampling regime in the current study was not

well suited to capturing measurements indicative of nitrogen biogeochemistry, or that our analytical capabilities were not sensitive enough to resolve nitrate variations in our samples.[110] The geochemistry data at Second Creek provides clear evidence, however, that Fe(III) and sulfate reduction are robust reactions in the hyporheic zone. The increased porewater phosphate concentrations could result from adsorbed phosphate released during reduction and dissolution of Fe(III) oxide and oxyhydroxide phases.[111–113] Remineralization of this organic matter can also increase phosphate[111,114] and methane,[2,115] which aligns with our observations of high analyte concentrations in the wetlands relative to the channel. Previous modeling at Second Creek indicates more reducing conditions in the wetland hyporheic zones because of their higher concentrations of organic carbon compared to in the channel.[52,53] Despite these concentration differences between wetland and stream channel sediments, the data indicate that cryptic S cycling linked to Fe and methane cycling is active in both environments.

Methane cycling and fluxes in the hyporheic zone

In addition to Fe and sulfate reduction, the highly reducing conditions in the hyporheic zone support methanogenesis, particularly in the organic-rich wetlands as described above. Indeed, a diversity of methanogenic archaea were previously discovered in relatively high abundances throughout the wetlands and stream channel sediments.[53] The rapid increase in CH_4 concentrations (Fig. 3) just below the depth where dissolved Fe(II) concentrations increase and SO_4^{2-} concentrations decrease shows that biogenic methanogenesis is spatially discrete from Fe(III) and sulfate reduction at most sampling times and locations. This is expected based on the lower redox potential of methanogenesis than for the other two redox reactions and which resulted in a deeper methanogenesis metabolic zone. Methane concentrations are generally lower in October than the other sampling months which also indicates that colder temperatures slightly suppressed methanogenesis, but there are still measurable amounts (near millimolar) in the wetlands (Fig. 3). Similar seasonal responses were previously observed in sediments and soils from a variety of environments resulting from decreased metabolic rates of methanogens due to lower temperatures and lower amounts of substrate availability, which also prompt shifts in the methanogen community composition and pathway (*e.g.*, aceticlastic *vs.* hydrogenotrophic).[116–119]

High CH_4 concentrations measured near the sediment–water interface during the June and July sampling times, however, suggests that methanogenesis occurred in very shallow sediments during that time. This has major implications for methane fluxes to the atmosphere, especially because the wetlands had either very shallow surface water or periods without inundation. Indeed, wetlands are the largest natural source of atmospheric methane, contributing ~20–39% of the global flux,[120,121] and rivers contribute another ~3% of the global emissions.[122] Upward hydraulic flux through the hyporheic zone during these months could contribute to the elevated levels in shallow sediments. However, the previous modeling studies at the site did not include methanogenesis at shallow depths, and they under-simulated CH_4 there.[52,53] This suggests that hydraulic flux cannot entirely produce the measured methane concentrations, and thus, some methanogenesis is likely co-occurring with Fe or sulfate reduction, as previously observed in other studies.[123–125] This is a particularly important consideration for implementing reactive transport models that prohibit co-occurring redox reactions that are contrary to thermodynamic ordering.

We suspect that the spuriously high CH_4 concentrations at various depths and times (*e.g.*, the channel sediment–water interface sample in October) resulted from trapping of methane bubbles in the porewater peepers. In dense, organic-rich sediments, CH_4-rich bubbles break out of sediments and travel through bubble tubes, escaping sediments at discrete, point-source locations in a process known as ebullition.[126–128] While sampling in the field, physical disruption of the wetland sediment from coring or from walking produced large amounts of bubbles, which could have contained CO_2 and CH_4. Upward hydrologic flux, diffusion, ebullition, and transport through arenchymous vascular plants (like those abundant in the wetlands) are all pathways of methane fluxes to the atmosphere,[129] and it is likely that all these processes play a role at various times at Second Creek.

The net flux of methane to the atmosphere from wetland sediments is controlled not only by methane production and transport processes, but also the amount of methane oxidized in the subsurface. Reactive transport modeling in previous studies[52,53] indicated that CH_4 was largely consumed in shallow sediments during upwelling hyporheic fluxes at Second Creek, possibly coupled to sulfate, Fe, and nitrate reduction. With the exception of June and July in the present study, methane levels are substantially lower in the upper 5 cm than at deeper sediment depths during many of the sampling points. This is particularly true in the cooler months and in the stream channel, which may reflect not only lower production rates at those depths, but also oxidation of methane that is transported upward. Evaluation of methane consumption from a wide range of freshwater wetlands demonstrates that a substantial, but highly variable, fraction of CH_4 generated in the subsurface, ranging from 1–90% of gross production, is consumed before reaching the atmosphere.[31] A long-held assumption is that most terrestrial CH_4 consumption occurs aerobically at the oxic–anoxic interface or near the rhizosphere, driven by methanotrophic bacteria that oxidize CH_4 for energy and biomass.[48,130,131] However, the anoxic conditions in the shallow reaches of the hyporheic zone at Second Creek during the warmer months with upwelling hydrologic conditions would have inhibited aerobic methane oxidation processes. Anaerobic methane oxidation (AOM) linked to abiotic Fe(III) reduction,[132–135] sulfate reduction,[39,40,136] or even humic acids,[132,137] would have been favorable in this environment. A metabolically active community of anaerobic methanotrophic (ANME) archaea was discovered in terrestrial and freshwater subsurface environments that couple AOM to sulfate reduction[138,139] and Fe reduction,[134] and these biotic processes could be coupled with abiotic processes that drive carbon cycling in the hyporheic zone. In these

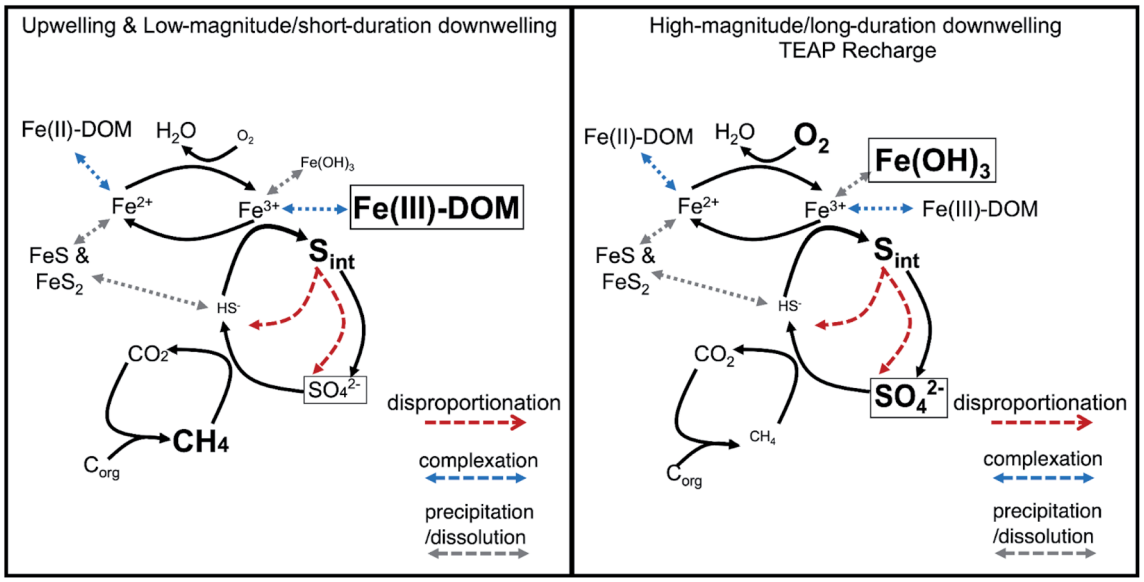

Fig. 10 Conceptual model of a proposed reaction network that sustains a cryptic S cycle linked to Fe reduction and methane oxidation in the Second Creek hyporheic zone under two different hydrologic regimes: upwelling and low-magnitude/short-duration downwelling (left panel) and High magnitude/long-duration downwelling (right panel). Solid black arrows indicate generalized abiotic and biotic reactions (*e.g.*, biological Fe(III) reduction could proceed *via* organic C oxidation instead of being coupled to sulfide oxidation to S intermediates). Species with boxes around them indicate important TEAs under the different conditions. Species with larger font size are expected to have increased concentrations in the system under the specific hydrologic conditions.

dynamic and heterogeneous environments in which TEAs are periodically recharged through hyporheic fluxes, the co-existence of methane redox cycling with S and Fe cycling appears to be a dominant process. This deviation from classical redox theory is notable and important for our understanding and prediction of ecosystem functioning. Supporting this finding, previous studies have identified functional overlaps and co-occurring bacterial metabolisms within similarly complex and heterogeneous riparian systems[140] and in coastal marine sediments.[141] The overall impact of this heterogeneity, the dynamic hyporheic exchange, and the biogeochemical conditions on methane fluxes to the atmosphere from the hyporheic zone deserve greater investigation.

Conclusions

The substantial contribution of organic and inorganic intermediate S compounds (*e.g.*, elemental sulfur, polysulfide, thiosulfate, sulfoxide, and thiophene) in the subsurface together with measurable Fe(III) oxyhydroxides and dissolved Fe(III)-complexes provide clear evidence of a cryptic S cycle coupled to Fe reduction in the Second Creek hyporheic zone. Importantly, the cryptic S cycle proceeds in stream and wetland sites and under both upwelling and downwelling conditions. Under dominantly upwelling conditions, including those with short-duration or low-magnitude shifts to downwelling, dissolved Fe(III)-complexes, as opposed to solid-phase Fe oxides, serve as TEAs for anaerobic sulfide oxidation that fuel the cryptic S cycle (Fig. 10, left panel). However, with high-magnitude or long-duration shifts to downwelling conditions, TEA recharge of O_2 and sulfate from surface water pushes the reaction network to utilize the newly formed electron acceptors (*e.g.* Fe oxide minerals) to fuel the cryptic S cycle (Fig. 10, right panel). Incorporating TEA variability, co-occurring redox reactions, and cryptic S cycling into future modeling efforts will be important for accurately capturing these complex and dynamic systems and thus improving predicted reaction rates and mass balances. In the future, further identifying and quantifying the primary S intermediate forms and developing genomically informed models will also be critical to expand our predictive understanding for how coupled biogeochemical cycles respond to environmental perturbations.[140]

At Second Creek, S recycling is likely fueled by high sulfate levels carried from the stream during downwelling hydrologic conditions and can have significant impacts on water quality. While some environmental sulfate loading has decreased with efforts to mitigate acid rain,[142,143] high sulfate concentrations from industrial and agricultural waters[38] remains problematic for many sensitive ecosystems, especially as sulfate is not federally regulated in all waste streams. Constructed wetlands have been used for the remediation of sulfate by promoting sulfate reduction to sulfide, and ultimately the precipitation of iron monosulfide (*e.g.*, FeS) and iron disulfide phases (*e.g.*, pyrite, chalcopyrite) that capture sulfide as well as heavy metals.[144–146] As can be seen from this study and others, however, there could be significant intermediate S species forming in wetlands that can be recycled to replenish sulfate within the system rather than precipitated in sulfide minerals. If the formation of these intermediate phases is correlated with sulfate enrichment in surface waters, it will be important to determine whether they represent a net S sink or potential sulfate source, which could have ramifications for long-term

water quality improvement in these wetlands. These intermediate S forms should be monitored more closely for ensuring the efficiency of these remediation approaches in controlling the fate and transport of contaminants in natural systems.

The fast recycling of S and Fe compounds in the hyporheic zone also has important implications for organic matter turnover, subsequently impacting CO_2 and CH_4 (greenhouse gas) emissions into the atmosphere. Some previous studies[147–150] suggest that increased sulfate, either from surface sources or regenerated in the subsurface through sulfide oxidation, will suppress methanogenesis and stimulate anaerobic methane oxidation, thereby mitigating greenhouse gas fluxes. With rapid climate change, as warming temperatures stimulate methanogenesis, much uncertainty remains as to how further interactions with Fe and S cycling will mediate impacts on net CH_4 fluxes, especially as hydrologic responses to intensified precipitation patterns will alter hyporheic zone mixing in unresolved ways. The complexity and heterogeneity of hyporheic zone hydrobiogeochemical interactions, such as those brought to light in this current study, necessitate further field-based examination. It is clear that advanced field methods, such as *in situ* geochemical sensors or real-time hydrologic flux monitoring for rapid-response sample collections, as well as higher spatial resolution geochemical analyses and isotopic analyses to capture reaction rates, would be beneficial to fully capture these transient and heterogeneous biogeochemical responses to dynamic hydrologic excursions. These additional methods, however, may still be unable to capture short-term hot moments/spots through ad hoc sampling approaches. Additionally, they can be expensive, not suitable for all environments, or require specialized analytical techniques not accessible to many (*e.g.*, electrochemical sensors). Regardless, our data point to the importance of several enigmatic sediment processes that can be incorporated into current and future modeling efforts[151–153] to accurately scale up and make predictions related to regional or global ecosystem, water quality, and climate impacts.

Data availability

Geochemistry data[154] is publicly available at the Environmental Data Initiative (EDI) data portal: **https://doi.org/10.6073/pasta/2a03c810518255ce147cccb10e92ace5**. Hydrologic data[155] is publicly available on Hydroshare at **https://doi.org/10.4211/hs.cf9090408d1f4b8193e8ce2b7e8ed313**.

Author contributions

JMT and CER contributed equally to this work and share first authorship of this manuscript. JMT and CER collected field samples, conducted geochemical analyses, performed data analysis and interpretation, and wrote sections of the manuscript. AJD and RS contributed to field sampling and geochemical analysis. KD and PAO conducted field preparation and field research. GHCN and CMS contributed to all aspects of the research and manuscript preparation and devised and supported the research.

Conflicts of interest

There are no conflicts to declare.

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