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# Reduction behaviors of tin oxides and oxyhydroxides during electrochemical reduction of carbon dioxide in an aqueous solution under neutral conditions†

Etsushi Tsuji, \*<sup>a</sup> Kaede Ohwan,<sup>a</sup> Tomoki Ishikawa,<sup>a</sup> Yuki Hirata,<sup>a</sup> Hiroyuki Okada, <sup>a</sup> Satoshi Suganuma <sup>ab</sup> and Naonobu Katada <sup>a</sup>

It is well known that tin oxides and oxyhydroxides show high selectivity for the electrochemical CO<sub>2</sub> reduction reaction (CO<sub>2</sub>RR) to form HCOOH in aqueous solutions. Tin oxides and oxyhydroxides are reduced to form metallic Sn during the CO<sub>2</sub>RR, and the formed interface between the oxide and metallic Sn plays important roles in the CO<sub>2</sub>RR. In this study, reduction behaviors of tin oxides and oxyhydroxide during the CO<sub>2</sub>RR were investigated. SnO, SnO<sub>2</sub> and tin oxyhydroxide containing both amorphous and crystalline phases were formed using solvothermal, sol-gel and precipitation methods, respectively. Reduction current densities of SnO<sub>2</sub> and the oxyhydroxide for the CO<sub>2</sub>RR and hydrogen evolution reaction at -0.8 V vs. RHE were higher than that of SnO, and the faradaic efficiency of the oxyhydroxides for formation of HCOOH and CO was >90%. Based on high-resolution TEM observation and EDS mappings, it was revealed that metallic Sn nanoparticles with a ~40 nm diameter were formed from SnO<sub>2</sub> and tin oxyhydroxides during the CO<sub>2</sub>RR via a dissolution and reductive deposition process. Aggregates of SnO<sub>2</sub> and the oxyhydroxide were dissolved in a neutral electrolyte solution during the CO<sub>2</sub>RR, and subsequently, metallic Sn nanoparticles with highly effective surface areas were formed on carbon electrodes via reductive deposition from dissolved Sn cations, leading to a higher reduction current. The thickness of native oxide layers formed on the surface of the metallic Sn particles in air after the CO<sub>2</sub>RR from the oxyhydroxide was greater than those of SnO and SnO<sub>2</sub>. Therefore, it is speculated that metallic surfaces of the former ones were more easily formed at the interface between SnO<sub>x</sub> and metallic Sn than those of the latter ones, leading to high selectivity for the CO<sub>2</sub>RR.

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## Sustainability spotlight

Our research aims to reveal reduction behaviors of Sn oxide-based catalysts, which are promising electrocatalysts for the CO<sub>2</sub> reduction reaction. Reduction behaviors of Sn oxide-based catalysts are important to improve catalytic activity for CO<sub>2</sub> reduction because the activity is enhanced by the interface between SnO<sub>x</sub> and metallic Sn formed through reduction of the catalysts during CO<sub>2</sub> reduction. Our research contributes toward achieving effective conversion of CO<sub>2</sub> into chemical feedstocks for a sustainable society. Our study emphasizes the importance of the following UN Sustainable Development Goals: affordable and clean energy (SDG 7), industry, innovation and infrastructure (SDG 9), and climate action (SDG 13).

## Introduction

It is necessary to decrease CO<sub>2</sub> concentration, which is the main cause of global warming on the earth.<sup>1</sup> The electrocatalytic CO<sub>2</sub> reduction reaction (CO<sub>2</sub>RR) in aqueous solutions using renewable energy, such as solar, wind and geothermal powers, is one

of the promising technologies to convert CO<sub>2</sub> into various chemical feedstocks,<sup>2,3</sup> for example, CO,<sup>4,5</sup> HCOOH,<sup>6–10</sup> CH<sub>4</sub>,<sup>11,12</sup> CH<sub>3</sub>OH<sup>13,14</sup> and lower olefins.<sup>14–18</sup> In electrocatalytic reactions, cathodic and anodic reactions are specifically separated, leading to facile collection of products. Meanwhile, H<sub>2</sub> is also produced as a by-product during the CO<sub>2</sub>RR in an aqueous solution. Therefore, it is important to develop electrocatalysts with both high activity and high selectivity for the CO<sub>2</sub>RR.

It is well known that tin-based electrocatalysts, such as metal,<sup>19</sup> alloys,<sup>20–22</sup> oxides,<sup>6–10</sup> oxyhydroxides,<sup>7–23</sup> and sulfides,<sup>24–26</sup> progress to form HCOOH via the CO<sub>2</sub>RR. In many cases, the electrocatalytic activity and selectivity of tin oxides for the CO<sub>2</sub>RR were higher than those of metallic Sn, although they

<sup>a</sup>Center for Research on Green Sustainable Chemistry, Tottori University, Tottori 680-8552, Japan. E-mail: tsuji@tottori-u.ac.jp

<sup>b</sup>Institute for Catalysis, Hokkaido University, Kita21, Nishi10, Kita-ku, Sapporo 001-0021, Japan

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were usually reduced to metallic Sn during the CO<sub>2</sub>RR because the metallic phase is the most thermodynamically stable under reduction conditions for the CO<sub>2</sub>RR.<sup>6,7,27</sup> Chen *et al.* reported that the electrocatalytic activity of Sn electrodes decreased upon the etching of the native oxide layer on the surface.<sup>6</sup> They also reported that CO<sub>2</sub>RR activity was enhanced through the electrodeposition of an oxide layer on the metallic surface, resulting in enhanced electrocatalytic activity for the CO<sub>2</sub>RR by the interface between metallic Sn and SnO<sub>x</sub>. Baruch *et al.* investigated the reduction behaviors of CO<sub>2</sub> on thin films of mixed Sn/SnO<sub>x</sub> species by *in situ* attenuated total reflectance infrared spectroscopy.<sup>7</sup> They proposed that the reduction of CO<sub>2</sub> was preceded by the two-electron reduction of SnO<sub>2</sub> to a Sn(II) hydroxide species, which can react with CO<sub>2</sub> to form surface-bound carbonate. Then, the carbonate species reacted with two electrons and a proton to form HCOOH, and returned to the Sn(II) hydroxide species by quick desorption of the formed HCOOH.

Some researchers investigated how to change the oxidation state of Sn atom in tin oxides during CO<sub>2</sub>RR.<sup>27–29</sup> Dutta *et al.* investigated the correlation between the oxidation state of SnO<sub>2</sub> and its electrocatalytic activity for CO<sub>2</sub>RR at various potentials using *operando* Raman spectroscopy.<sup>28</sup> SnO<sub>2</sub> showed a high faradaic efficiency to generate HCOOH at a moderate cathodic potential, where SnO<sub>2</sub> is thermodynamically unstable. Meanwhile, the efficiency to generate HCOOH was significantly decreased at very negative potentials, where the oxide is reduced to metallic Sn. Ning *et al.* reported that SnO<sub>2</sub> grown onto a carbon cloth by a hydrothermal method was partially reduced to Sn from the bottom of SnO<sub>2</sub>, as evidenced in by XRD, *in situ* Raman spectroscopy and XPS measurement.<sup>27</sup> DFT calculations show that the formation of an interface between SnO<sub>2</sub> and metallic Sn enhanced the stability of the adsorbed HCOO<sup>−</sup> intermediate, leading to the highly selective production of HCOOH. On the other hand, there are very few reports on the difference of the reduction behaviors of tin oxide and oxyhydroxide catalysts during CO<sub>2</sub>RR.

In this study, we formed SnO, SnO<sub>2</sub> and a tin oxyhydroxide containing amorphous and crystalline Sn<sub>3</sub>O<sub>2</sub>(OH)<sub>2</sub> by different methods, and investigated their reduction behaviors by XRD, SEM, high-resolution TEM, STEM, and EDS mapping before and after CO<sub>2</sub>RR. We found that SnO was reduced to metallic Sn while maintaining its morphology during CO<sub>2</sub>RR; whereas SnO<sub>2</sub> and the tin oxyhydroxide were dissolved into a neutral electrolyte solution, and then formed metallic Sn nanoparticles of ~40 nm diameter on the electrodes by reductive deposition from dissolved Sn cations.

## Experimental

### Preparation of catalysts

In this study, three types of tin oxides or oxyhydroxides were prepared by different methods; SG-, P- and ST-SNO were prepared by sol-gel, precipitation, and solvothermal methods, respectively.

In the case of SG-SNO, SnCl<sub>2</sub>·2H<sub>2</sub>O (Fujifilm Wako Pure Chemical Corp. 97.0%) and citric acid monohydrate (Fujifilm

Wako Pure Chemical Corp. 99.5%) were dissolved in ion-exchanged water. The molar ratio of SnCl<sub>2</sub>/(citric acid) was 0.5. The solution was stirred at 363 K for polymerization. The obtained gel was pre-calcined at 523 K for 1 h in air, ground for 30 min in a mortar, and then calcined at 723 K for 6 h in air.

In the case of P-SNO, SnCl<sub>2</sub>·2H<sub>2</sub>O was a dissolved mixed solution of ion-exchanged water and propylene glycol (Fujifilm Wako Pure Chemical Corp. 99.0%) (1 : 1 volume ratio). Diethanolamine (Tokyo Kasei Corp., 99.0%) was added to the obtained solution drop-by-drop. The obtained white solid was collected by centrifugation at 4000 rpm, and dried at 383 K for overnight.

In the case of ST-SNO, the P-SNO suspension before the centrifugation shown above was put into a batch autoclave reactor, and heated at 433 K for 3 h without stirring. After the reaction, the suspension was cooled with water, and the solid was collected by centrifugation at 4000 rpm and dried overnight at 383 K.

### Characterization

The crystallinities of the as-prepared ST-, SG- and P-SNO powders and the carbon paper (CP) electrodes loading ST-, SG- and P-SNO as the working electrodes before and after CO<sub>2</sub>RR (the details of preparation of the working electrodes for XRD are described in the “Preparation of electrodes and catalysts before and after CO<sub>2</sub>RR for analysis” section) were investigated by powder X-ray diffraction (XRD; Rigaku, Ultima IV) using Cu K $\alpha$  radiation with some pieces of small Cu foil as a reference. The morphology of the electrodes before and after CO<sub>2</sub>RR was investigated using a scanning electron microscope (SEM; HITACHI S-4800) operated at 5 kV. The oxidation states of the working electrodes before and after CO<sub>2</sub>RR were investigated by Sn 3d and O 2p X-ray photoelectron spectroscopy (XPS; JEOL JPS-9200) using Mg K $\alpha$  radiation. The binding energy was calibrated using the C 1s peak at 285.2 eV (included contamination). The local morphology, crystallinity and chemical composition of each catalyst particle were investigated by Cs-corrected scanning and transmission electron microscope with EDS (S/TEM-EDS; Titan3 G2 60-300, FEI Co.) at 300 kV. Each electrode and sample preparation after CO<sub>2</sub>RR are shown in the section, “Preparation of electrodes and catalysts before and after CO<sub>2</sub>RR for Analysis”.

### Preparation of catalyst inks and electrodes

A catalyst ink was prepared as follows. ST-SNO, SG-SNO or P-SNO powder (48 mg), a commercial acetylene carbon black (AB, 32 mg, STREM Chemicals, Inc.) as the conductive support, and 5 wt% Nafion solution (160 mm<sup>3</sup>, FUJIFILM Wako Pure Chemical Corp., DE521 CS type) were dispersed into a mixed solution of ethanol and ion-exchanged water (ethanol : water volume ratio = 1 : 3, total volume of the mixture was 5.0 cm<sup>3</sup>). The suspension was sonicated for at least 1 h to obtain a homogenous catalyst ink. The working electrodes were prepared by drop-casting the catalyst inks onto carbon paper electrodes (CP, Toyo Tanso Co., PERMA-FOIL, 2.0 × 1.0 cm<sup>2</sup>). The amount of ink was adjusted to maintain a concentration of



$1.0 \text{ mg}_{\text{cat}} \text{ cm}^{-2}$  of the catalyst loading on the apparent electrode surface. The lateral and back sides of the carbon paper were covered with epoxy resin for insulation (effective area is  $1.0 \times 1.0 \text{ cm}^2$ ).

### Electrochemical activity and selectivity for CO<sub>2</sub>RR

The CO<sub>2</sub>RR activity was investigated by electrochemical measurements using a normal three-electrode H-type cell with an anion exchange membrane (AGC Engineering Co., Ltd, AMVN). ST-SNO, SG-SNO or P-SNO coated with carbon paper was used as the working electrode. A platinum foil ( $1.5 \times 1.5 \text{ cm}^2$ ) and Ag/AgCl/KCl (satd.) were used as the counter and reference electrodes, respectively.  $0.1 \text{ mol dm}^{-3} \text{ NaHCO}_3$  solution saturated with CO<sub>2</sub> (pH 7) was used as an electrolyte solution. The cyclic voltammograms (CVs) were obtained with a commercial potentiostat and potential programmer (Toho Technical Research Co., Ltd, Model PS-04) with a potential sweep rate of  $10 \text{ mV s}^{-1}$  and a constant stirring rate at 350 rpm for three cycles. The potential  $U$  was converted from the Ag/AgCl/KCl (satd.) reference scale to RHE using the following equation.

$$U \text{ V vs. RHE} = U \text{ V vs. Ag/AgCl/KCl (satd.)} + 0.199 \text{ V} + 0.059 \text{ V} \times \text{pH}$$

Constant potential measurements were carried out at  $-0.8 \text{ V}$  vs. RHE for 6 h. After applying potential for 6 h, the generation of the gas component was corrected by  $1.0 \text{ cm}^3$  of a gas-tight syringe, and analyzed by a gas chromatograph with a thermal conductivity detector (Shimadzu Corp., GC-8A) and an MS-5A column through argon carrier gas. The generation of the liquid component was corrected by  $50 \text{ mm}^{-3}$  of a syringe, and analyzed by a high-performance liquid chromatograph with a UV detector (JASCO Corp., PU4580 and UV4075) and packed columns (Rspak KC-811 and Unifinepak C18). The faradaic efficiencies of CO, HCOOH and H<sub>2</sub> ( $FE_{\text{CO}}$ ,  $FE_{\text{HCOOH}}$  and  $FE_{\text{H}_2}$ ) were calculated from the ratio of electrons used in each reaction and that of the total electrons.

### Preparation of electrodes and catalysts before and after CO<sub>2</sub>RR for analysis

Each ST-, SG- and P-SNO/CP electrode was evaluated by XRD after CVs were performed between  $0.5$  and  $-1.2 \text{ V}$  vs. RHE with a potential sweep rate of  $10 \text{ mV s}^{-1}$  for 3 cycles, and following constant potential measurements at  $-0.8 \text{ V}$  vs. RHE for 0.25, 0.5, 1 and 2 h. After each electrochemical measurement, the CP sheets loaded with ST-, SG- and P-SNO were rinsed with ion-exchanged water and dried in an air stream. Subsequently, the samples were investigated by XRD. Those after-constant potential measurements at  $-0.8 \text{ V}$  vs. RHE for 2 h were also investigated by SEM and XPS. The ST-, SG- and P-SNO powders were scratched from the surface of the electrodes after constant potential measurement at  $-0.8 \text{ V}$  vs. RHE for 2 h, and then investigated by high-resolution (S)TEM and EDS mapping.

## Results and discussion

### Electrocatalytic activity for CO<sub>2</sub>RR

Fig. 1 shows the XRD patterns of the samples before the electrochemical reaction. ST- and SG-SNO showed XRD patterns that were attributed to tetragonal SnO and rutile SnO<sub>2</sub>, respectively. Conversely, P-SNO shows both a broad halo pattern between 30 and 50 degree and a weak pattern attributed to crystalline Sn<sub>3</sub>O<sub>2</sub>(OH)<sub>2</sub>, indicating that it was a mixture of the amorphous and crystalline oxyhydroxide phases. Therefore, ST-, SG- and P-SNO were tetragonal SnO, rutile SnO<sub>2</sub> and a mixture of crystalline Sn<sub>3</sub>O<sub>2</sub>(OH)<sub>2</sub> and amorphous phase, respectively.

Fig. 2 shows the electrocatalytic activities of ST-, SG- and P-SNO, together with that of a metallic Sn electrode formed by electrodeposition (details of the electrodeposition is shown in the ESI†). From the CV measurements shown in Fig. 2a, ST-, SG- and P-SNO showed large redox peaks attributed to the reduction and oxidation of the Sn cation/Sn metal at  $-0.6$ – $0.4 \text{ V}$  vs. RHE, whereas the electrodeposited Sn electrode showed small ones (black dash line in the inset figure). In all samples, the reduction current densities were observed at approximately below  $-0.6 \text{ V}$  vs. RHE. The onset potentials and reduction current densities of ST-, SG- and P-SNO were higher than those of metallic Sn, suggesting that the electrocatalytic activities for the reduction reactions on the oxides and oxyhydroxide were higher than that of metallic Sn. From time courses of the current density at  $-0.8 \text{ V}$  vs. RHE (Fig. 2b), the reduction current densities of ST-SNO and metallic Sn showed almost constant values from the initial stage of CO<sub>2</sub>RR, whereas those of SG- and P-SNO changed at the initial stage and became constant after 1 h. The time required for the current density to stabilize was observed to be shorter for SG-SNO compared to P-SNO. The saturated current density of SG-SNO at more than 2 h was

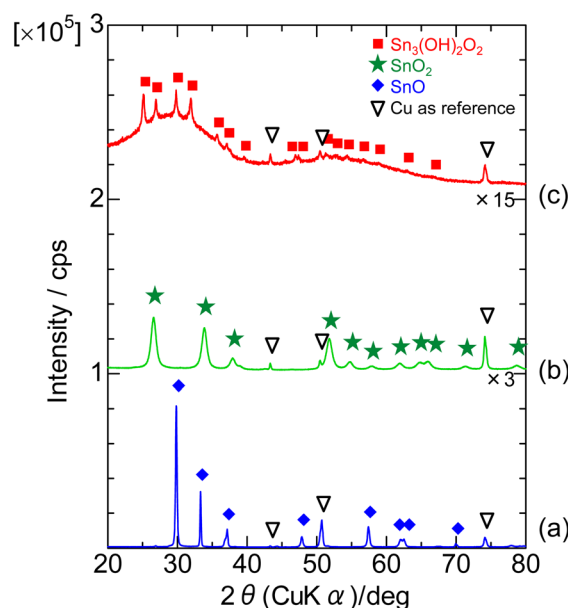


Fig. 1 XRD patterns of (a) ST-SNO, (b) SG-SNO and (c) P-SNO powders before the electrochemical reaction.



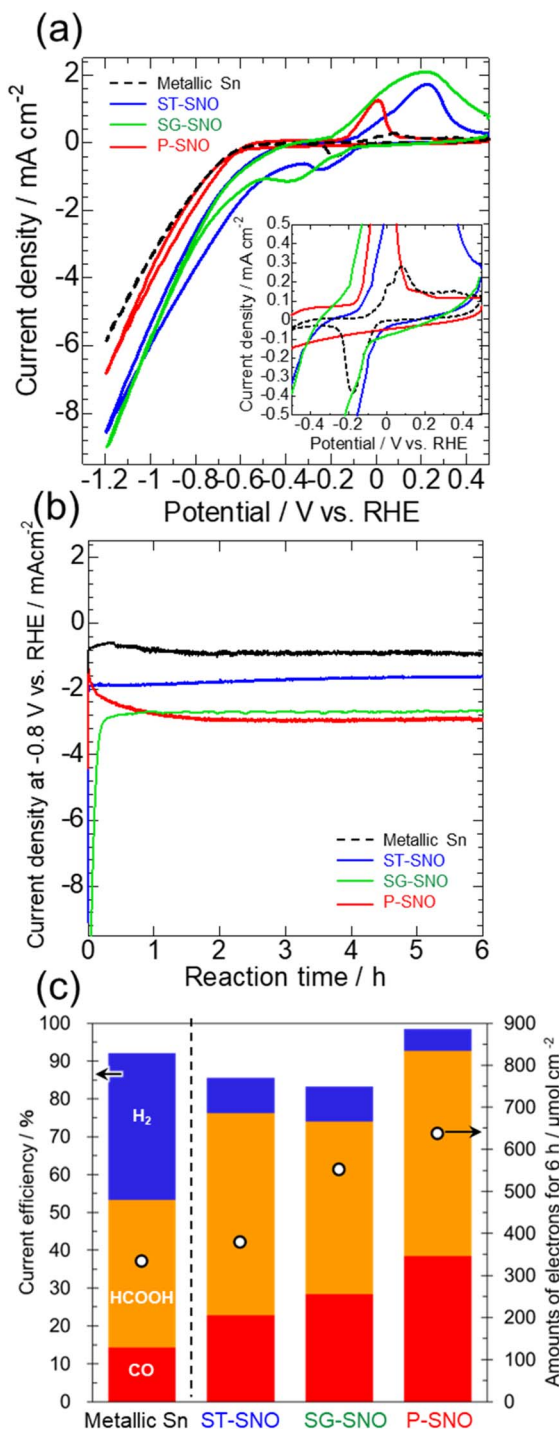


Fig. 2 (a) Cyclic voltammograms, (b) time courses of current densities at  $-0.8$  V vs. RHE for 6 h, and (c) the amounts of electrons used for the reduction reaction and current efficiency to form each product at  $-0.8$  V vs. RHE for 6 h of ST-, SG- and P-SNO measured in  $\text{CO}_2$ -saturated  $\text{NaHCO}_3$  aqueous solution, together with those of a metallic Sn formed by the electrochemical deposition. The inset figure in (a) is the enlarged cyclic voltammograms between  $-0.5$  and  $0.5$  V vs. RHE.

comparable to that of P-SNO. Fig. 2c shows the amounts of electrons used in the reduction reactions, and the faradaic efficiency of all samples at  $-0.8$  V vs. RHE for 6 h. The amounts of electrons involved in the reduction reactions on SG-SNO and

P-SNO were clearly greater than that on ST-SNO and metallic Sn. Meanwhile, the faradaic efficiencies for  $\text{CO}_2\text{RR}$  ( $\text{FE}_{\text{CO}} + \text{FE}_{\text{HCOOH}}$ ) on ST-, SG- and P-SNO were much higher than that on metallic Sn. Furthermore,  $\text{FE}_{\text{H}_2}$  on P-SNO was lowest among them, and the total of  $\text{FE}_{\text{CO}}$  and  $\text{FE}_{\text{HCOOH}}$  was  $>90\%$ . The  $\text{FE}_{\text{HCOOH}}$  on P-SNO was comparable to that on ST-SNO ( $\sim 55\%$ ), and greater than those of SG-SNO and metallic Sn. These results were in agreement with the previous report.<sup>7</sup>

### Change of the crystallinity, morphology and chemical state of catalysts after $\text{CO}_2\text{RR}$

It is well known that tin oxides are reduced during  $\text{CO}_2\text{RR}$ .<sup>6,7,27</sup> Thus, methods for how to change ST-, SG- and P-SNO during  $\text{CO}_2\text{RR}$  in various conditions were investigated. Fig. 3 shows the XRD patterns of all samples after 3 cycles of CV between  $0.5$  V to  $1.2$  V vs. RHE with a potential sweep rate of  $10$  mV s<sup>-1</sup> and following constant potential measurements at  $-0.8$  V vs. RHE for 0.25, 0.5, 1 and 2 h, respectively. In the case of ST-SNO, the strong pattern of metallic Sn was detected just after CV cycles with a small pattern from the retained tetragonal SnO. After more than 0.25 h of constant potential measurements, the pattern of the tetragonal SnO completely disappeared. These results indicated that ST-SNO, which is tetragonal SnO, was easily reduced to metallic Sn during  $\text{CO}_2\text{RR}$ . Conversely, in the case of SG-SNO and P-SNO, only very weak patterns of metallic Sn were observed after CV cycles, together with that of SnO<sub>2</sub>. Then, the intensity of the metallic Sn pattern increased with increasing applied potential time, even though that of SG-SNO increased more rapidly than that of P-SNO.

Fig. 4 shows the SEM observations of ST-, SG- and P-SNO before and after  $\text{CO}_2\text{RR}$  at  $-0.8$  V vs. RHE for 2 h. ST-SNO comprised cubic shape particles with sizes in the micron to sub-micron scales before  $\text{CO}_2\text{RR}$  (Fig. 4a-1), and such cubic shapes were almost retained after the reactions (Fig. 4a-2). Conversely, SG- and P-SNO were aggregates of several micrometers in size surrounded by AB particles before  $\text{CO}_2\text{RR}$  (Fig. 4b-1 and c-1). The primary particle of SG-SNO was a round shape with a diameter of 20–30 nm (inset of Fig. 4b-1), whereas that of P-SNO was an uneven shape that was smaller than 10 nm (inset of Fig. 4c-1). In both cases of SG- and P-SNO after  $\text{CO}_2\text{RR}$ , similar aggregates were observed, and some holes were observed in AB layers (Fig. 4b-2 and c-2).

To further investigate the details of the morphologies and chemical compositions of the samples after  $\text{CO}_2\text{RR}$  for 2 h, at which point the current density of all samples became constant, they were characterized by high-resolution TEM and EDS mappings. From the high-resolution TEM images, the cubic shape of ST-SNO was almost retained even after  $\text{CO}_2\text{RR}$  (Fig. 5a-1), as mentioned in the SEM image shown in Fig. 4a-2. Conversely, both SG-SNO and P-SNO were core-shell nanoparticles with diameters of 40 nm (Fig. 5b-1 and c-1), even though the shape of SG-SNO before  $\text{CO}_2\text{RR}$  was different from that of P-SNO, which was a mixture of crystalline nanoparticles with diameters of about 5 nm and amorphous grains (Fig. S1†). The layer thickness of both shells was about 5 nm. The lattice fringes, which were attributed to the (200) face of metallic Sn, were observed in the inner parts of



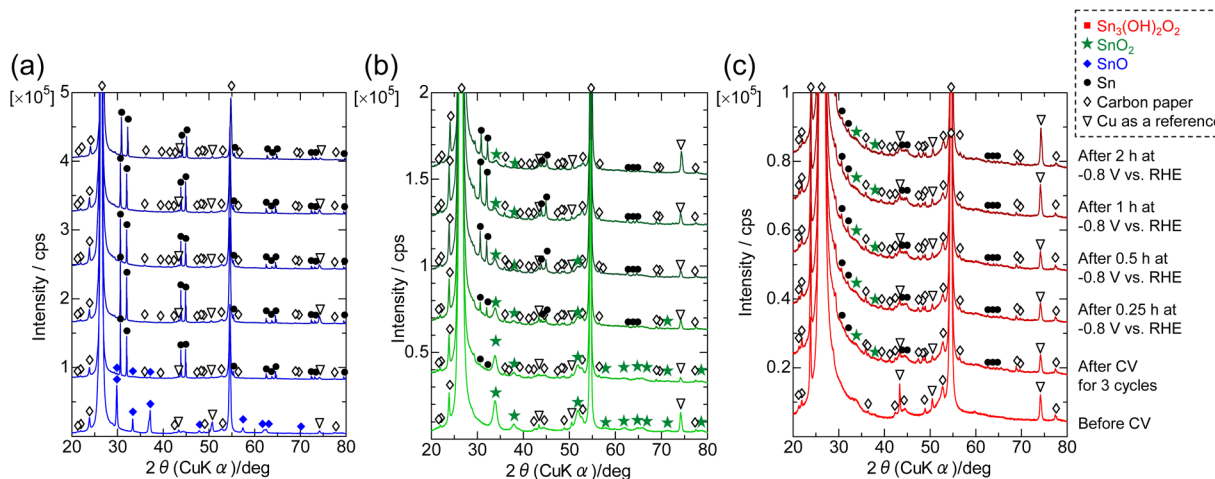


Fig. 3 XRD patterns of (a) ST-SNO, (b) SG-SNO and (c) P-SNO loaded on carbon paper electrodes before and after CV between 0.5 and  $-1.2$  V vs. RHE with  $10 \text{ mV s}^{-1}$  for 3 cycles, and constant potential measurements at  $-0.8$  V vs. RHE for 0–2 h in  $\text{CO}_2$  saturated  $\text{NaHCO}_3$  aqueous solution.

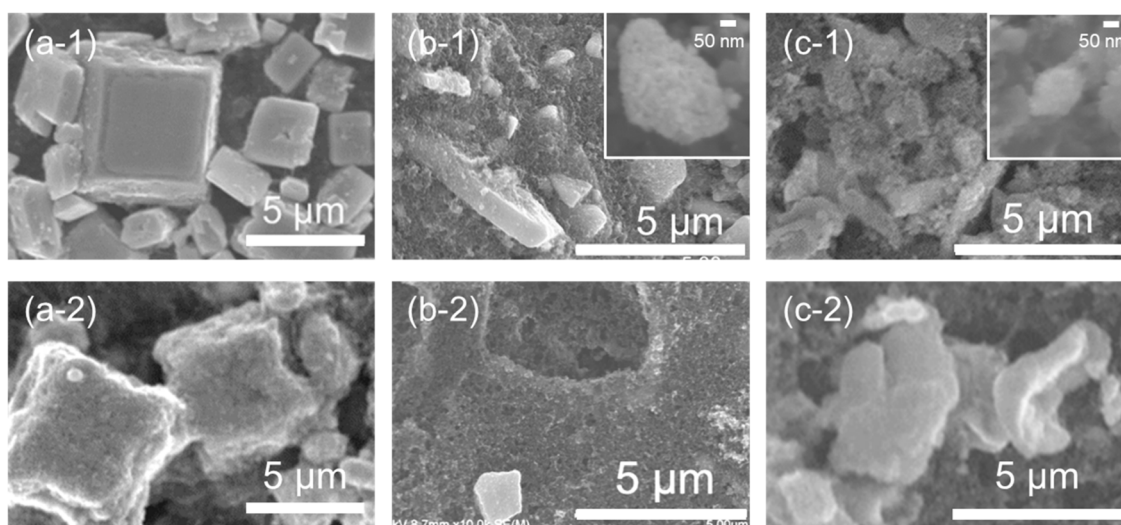


Fig. 4 SEM images of (a) ST-SNO, (b) SG-SNO and (c) P-SNO/carbon paper electrodes (1) before and (2) after constant potential measurements at  $-0.8$  V vs. RHE for 2 h in  $\text{CO}_2$  saturated  $\text{NaHCO}_3$  aqueous solution. Inset figures in (b-1) and (c-1) are the enlarged figures of SG-SNO and P-SNO, respectively, before the reaction.

SG-SNO and P-SNO, whereas no lattice fringe was observed in the outer parts, suggesting that the outer parts were amorphous (Fig. 5b-2 and c-2). From EDS mappings (Fig. 5a-3–c-3), the inner parts of all samples after  $\text{CO}_2\text{RR}$  were metallic Sn. Meanwhile, the outer parts, which may be oxidized in air after  $\text{CO}_2\text{RR}$ , were tin oxide. In all samples, the molar ratio of Sn/O of the outer oxide layer was about 1, suggesting that the average oxidation state of Sn in the outer layer was divalent.

Fig. 6 shows the XPS spectra of ST-, SG- and P-SNO after  $\text{CO}_2\text{RR}$ . From the Sn 3d spectra. All samples contained metallic and cationic Sn species after  $\text{CO}_2\text{RR}$ , whereas those samples contained cationic Sn before  $\text{CO}_2\text{RR}$  (Fig. S2†). These results were consistent with the HR-TEM and EDS mappings, as shown in Fig. 5. The molar ratios of  $\text{Sn}^{2+}/\text{Sn}^0$  at ST-, SG- and P-SNO were 8.3, 7.7 and 16.7, suggesting that the thickness of the

outer layer of P-SNO was greater than those of ST-SNO and SG-SNO. Conversely, from the O 2p spectra, all spectra were deconvoluted to four peaks assigned to the lattice O, OH species, adsorbed  $\text{H}_2\text{O}$ , and oxygen of Nafion. The molar ratios of the OH species to the lattice oxygen of ST-, SG- and P-SNO were 0.35, 0.44 and 0.39, respectively. These results suggest that the amount of OH species containing in the oxide layer, which is speculated to form by oxidation of the metallic Sn surfaces in air, and was almost comparable among all samples.

#### Reduction behaviors of tin oxides and oxyhydroxides during $\text{CO}_2\text{RR}$

Fig. 7 shows a schematic illustration of how to reduce ST-, SG- and P-SNO to metallic Sn. In the case of ST-SNO, which is



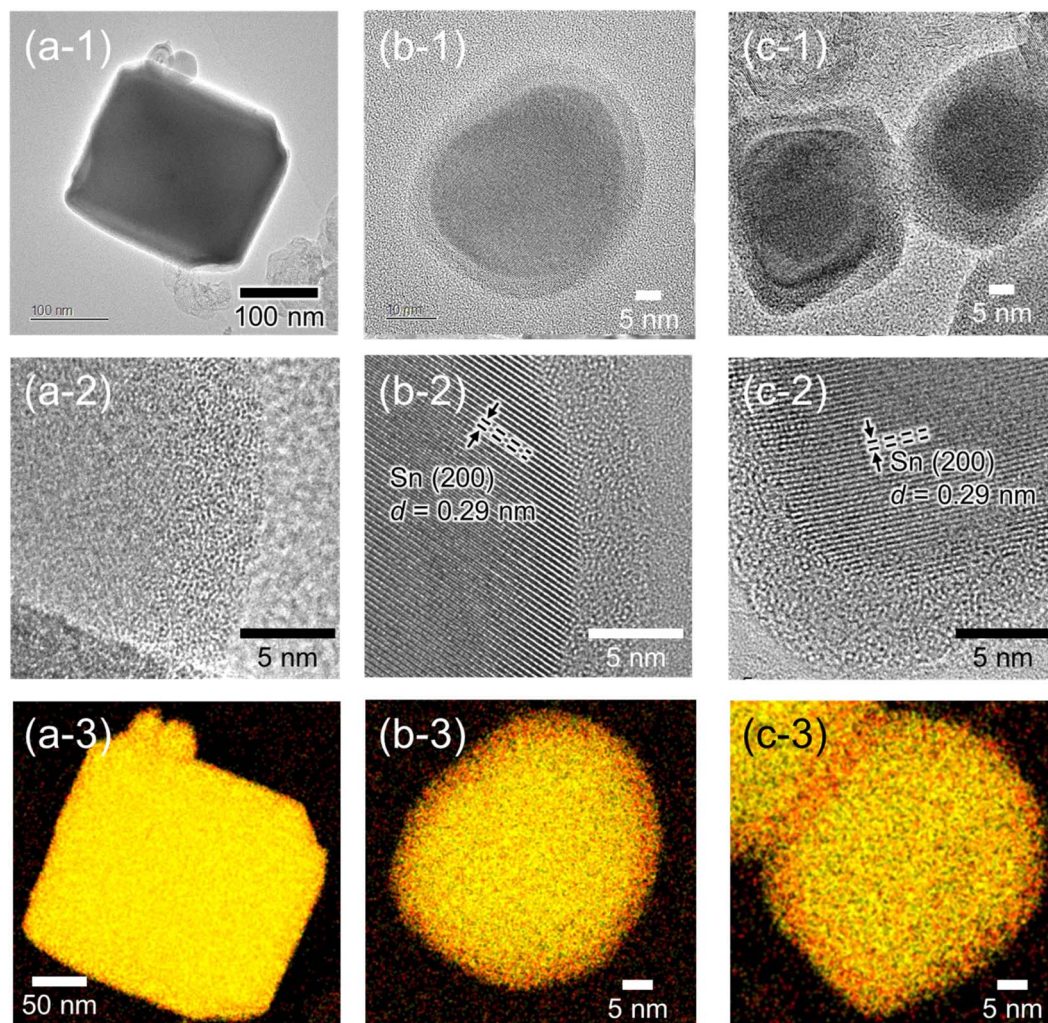


Fig. 5 (1 and 2) HR-TEM images and (3) EDS mappings of (a) ST-SNO, (b) SG-SNO and (c) P-SNO/carbon paper electrodes after constant potential measurements at  $-0.8$  V vs. RHE for 0–2 h in  $\text{CO}_2$  saturated  $\text{NaHCO}_3$  aqueous solution.

a tetragonal SnO formed by the solvothermal method, SnO was rapidly reduced to metallic Sn during  $\text{CO}_2\text{RR}$  (Fig. 3a) and the cubic shape of SnO was almost maintained after  $\text{CO}_2\text{RR}$  (Fig. 4a and 5a). It is well known that the SnO phase is thermodynamically more unstable than the  $\text{SnO}_2$  phase. In addition, from chemical composition analysis, ST-SNO had some oxygen vacancies and the chemical composition of ST-SNO was  $\text{SnO}_{0.81}$ . Therefore, it is speculated that the unstable phase induced by  $\text{Sn}^{2+}$  and the oxygen vacancies promote the rapid reduction itself to the metallic phase during  $\text{CO}_2\text{RR}$ . These results indicate that the tetragonal SnO phase was directly reduced to metallic Sn phase at  $-0.8$  V vs. RHE (Fig. 7a). Conversely, in the case of both SG-SNO and P-SNO, which were  $\text{SnO}_2$  and a mixture of amorphous tin oxyhydroxide and crystalline  $\text{Sn}_3\text{O}_2(\text{OH})_2$ , respectively, they were gradually reduced to the metallic Sn phase (Fig. 3b and c). Furthermore, both particles had a round shape with the diameter of  $\sim 40$  nm, even though the shape of  $\text{SnO}_2$  was different from that of the amorphous oxyhydroxide before  $\text{CO}_2\text{RR}$  (Fig. 4b, c and 5b, c). In addition, when ultrasmall particles of  $\text{SnO}_2$  with diameters of less than 5 nm (Fig. S3†)

synthesized by reverse micelle deposition<sup>30</sup> and solvothermal method (the details are shown in the ESI†) were reduced at  $-0.8$  V vs. RHE for 2 h, similar round shape particles were formed (Fig. S4†). These results suggest that almost the same round shape metallic Sn particles were formed on  $\text{SnO}_2$  during  $\text{CO}_2\text{RR}$  by a different reduction mechanism from SnO, which almost maintained the shape of the particles even after  $\text{CO}_2\text{RR}$ . It is speculated that  $\text{SnO}_2$  and oxyhydroxides were first dissolved into the electrolyte solution at pH 7 during  $\text{CO}_2\text{RR}$ . Then, the Sn cation was reduced to the metallic phase and formed round shape nanoparticles on the electrodes.

These findings open up a new way towards improving the electrocatalytic activity of tin oxides and oxyhydroxides for  $\text{CO}_2\text{RR}$ . For example, improving the activity by the formation of nanostructures, such as nanoparticles, nanofibers, and nano-sheet, is maybe more useful for tetragonal SnO, not  $\text{SnO}_2$  and tin oxyhydroxides. It is also important to investigate the reduction behaviors to improve the catalytic activity of other metal oxides and oxyhydroxides catalysts for  $\text{CO}_2\text{RR}$ .



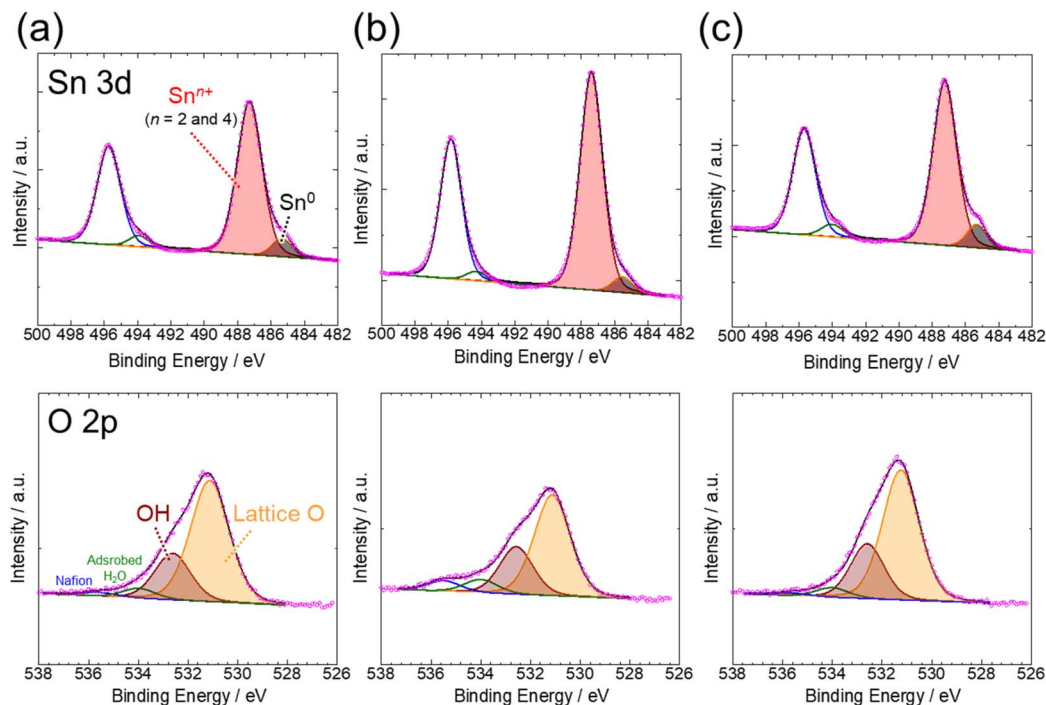


Fig. 6 XPS spectra of Sn3d and O2p of (a) ST-SNO, (b) SG-SNO and (c) P-SNO/carbon paper electrodes after constant potential measurements at  $-0.8$  V vs. RHE for 0–2 h in  $\text{CO}_2$  saturated  $\text{NaHCO}_3$  aqueous solution.

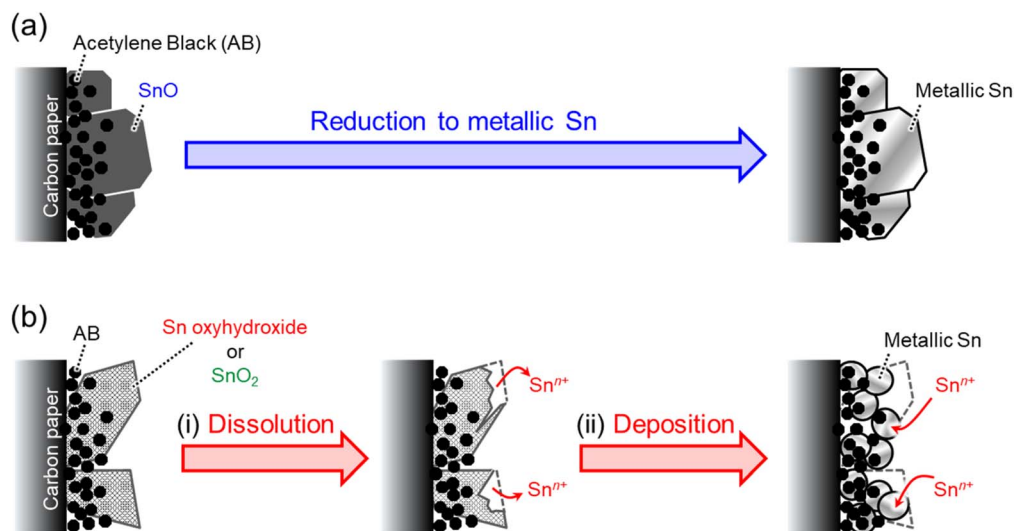


Fig. 7 Schematics of the reduction behaviors of (a) SnO and (b) tin oxyhydroxide or  $\text{SnO}_2$  during the  $\text{CO}_2\text{RR}$ .

### Correlation between reduced catalysts and electrocatalytic activity for $\text{CO}_2\text{RR}$

The time courses of the current densities of ST-, SG- and P-SNO are consistent with their reduction rates. ST-SNO was rapidly reduced to metallic Sn during  $\text{CO}_2\text{RR}$  (Fig. 3a) without changing the morphology (Fig. 4a). Consequently, the current density was constant at  $-0.8$  V vs. RHE despite the initial stage of  $\text{CO}_2\text{RR}$  (Fig. 2b). In contrast, SG- and P-SNO were dissolved and reduced to the metallic phase, and the reduction rate of SG-SNO was faster than that of P-SNO (Fig. 3b and c). The current densities of both

SG- and P-SNO changed during the initial stage of  $\text{CO}_2\text{RR}$  (Fig. 2b). The time required for the current density to stabilize was shorter for SG-SNO ( $\sim 0.25$  h) compared to P-SNO ( $\sim 2$  h), reflecting the relatively fast reduction rate of SG-SNO. As mentioned above, the current densities of SG- and P-SNO changed during the initial stage of  $\text{CO}_2\text{RR}$ , whereas that of ST-SNO was nearly constant (Fig. 2b). This is because the morphology of the SG- and P-SNO changed during  $\text{CO}_2\text{RR}$  (Fig. 4b, c and 5b, c), leading to an increase or decrease of the effective surface areas (ECSA) of the electrocatalysts. Conversely,



ST-SNO maintained its morphology due to the rapid reduction to Sn metal (Fig. 4a and 5a). The current density and total amount of electrons used in CO<sub>2</sub>RR and HER of P-SNO was almost the same as those of SG-SNO (Fig. 2b and c) and higher than that of ST-SNO. One of the possible reasons is that the particle size of SG- and P-SNO after reduction was almost the same and smaller than that of ST-SNO, leading to a higher ECSA of both electrocatalysts.

Selectivity of P-SNO for CO<sub>2</sub>RR was higher than those of the others. It is well known that surface hydroxyls play an important role due to the formation of an intermediate, such as \*OCHO by not only experimental analysis such as ATR-IR<sup>8,23</sup> but also DFT calculations.<sup>31,32</sup> From O 2p XPS spectra of the samples after CO<sub>2</sub>RR shown in Fig. 6, the amounts of hydroxyls of P-SNO were not clearly higher than those of ST- and SG-SNO. However, molar ratio of Sn<sup>+/0</sup> of the former one was higher than those of the latter ones (Sn 3d spectra shown in Fig. 6), suggesting that the native oxide layer formed on the metallic Sn nanoparticles at P-SNO is thicker than those of ST- and SG-SNO. These results suggested that the surface of the metallic Sn nanoparticles formed from P-SNO was more easily oxidized than those from ST- and SG-SNO, leading to the rapid formation of an interface between SnO<sub>x</sub> and metallic Sn, which acts as the active site for CO<sub>2</sub>RR.<sup>6,27</sup> Additionally, selectivity for HCOOH formation at −0.8 V vs. RHE on ST- and P-SNO, which were mainly composed of Sn<sub>3</sub>O<sub>2</sub>(OH)<sub>2</sub> and SnO, were higher than that of SG-SNO, which consisted of SnO<sub>2</sub> (Fig. 2c). In the previous report, Baruh *et al.* suggested that the Sn<sup>2+</sup> species on Sn showed higher activity for HCOOH formation compared to Sn<sup>4+</sup> on Sn. Furthermore, it is difficult for Sn<sup>4+</sup> to reduce to Sn<sup>2+</sup> at −0.788 V vs. RHE.<sup>7</sup> Therefore, it is speculated that the amount of Sn<sup>2+</sup> species in ST- and P-SNO at −0.8 V vs. RHE was higher than that in SG-SNO, leading to higher selectivity for HCOOH formation.

## Conclusions

SnO, SnO<sub>2</sub> and a mixture of amorphous tin oxyhydroxide and a small amount of crystalline Sn<sub>3</sub>O<sub>2</sub>(OH)<sub>2</sub> were formed by various methods. SnO<sub>2</sub> and the oxyhydroxides showed higher reduction current density than SnO. Furthermore, the faradaic efficiency of the oxyhydroxide for the formation of HCOOH and CO was more than 90%. SnO was reduced to metallic Sn particles, while maintaining the sub-micron scale cubic shape during CO<sub>2</sub>RR. Conversely, both SnO<sub>2</sub> and the oxyhydroxide were dissolved into the electric solution at pH 7, and reduced and electrodeposited as metallic Sn nanoparticles with a diameter of ~40 nm on the electrode. Therefore, highly effective surface areas of metallic Sn nanoparticles were formed from SnO<sub>2</sub> and the oxyhydroxide, which was induced to higher current density than SnO. On the other hand, it is speculated that the oxide layer was easily formed on the surface of metallic Sn from the oxyhydroxide compared to SnO and SnO<sub>2</sub>, leading to higher selectivity for CO<sub>2</sub>RR.

## Data availability

All the data supporting this article have been included in the manuscript in the form of figures and part of the ESI.†

## Author contributions

E. Tsuji was a supervisor, and contributed to the concept of this study, and writing the manuscript. K. Ohwan, T. Ishikawa, Y. Hirata and H. Okada measured and analyzed the data, and revised the manuscript. S. Sukanuma and N. Katada were co-supervisors, and discussed the data and reviewed the manuscript.

## Conflicts of interest

There are no conflicts to declare.

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