

Cite this: *Chem. Sci.*, 2025, 16, 9010

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A new type of C₂H₂ binding site in a *cis*-bridging hexafluorosilicate ultramicroporous material that offers trace C₂H₂ capture†

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Hybrid ultramicroporous materials (HUMs) comprising hexafluorosilicate (SiF₆²⁻, SIFSIX) and their variants are promising physisorbents for trace acetylene (C₂H₂) capture and separation, where the inorganic anions serve as *trans*-bridging pillars. Herein, for the first time, we report a strategy of fluorine binding engineering in these HUMs *via* switching the coordination mode of SIFSIX from traditional *trans* to rarely explored *cis*. The first example of a rigid HUM involving *cis*-bridging SIFSIX, SIFSIX-bidmb-Cu (bidmb = 1,4-bis(1-imidazolyl)-2,5-dimethylbenzene), is reported. The resulting self-interpenetrated network is found to be water stable and exhibits strong binding to C₂H₂ but weak binding to C₂H₄ and CO₂, affording a high Q_{st} of 55.7 kJ mol⁻¹ for C₂H₂, a high C₂H₂ uptake of 1.86 mmol g⁻¹ at 0.01 bar and high ΔQ_{st} values. Breakthrough experiments comprehensively demonstrate that SIFSIX-bidmb-Cu can efficiently capture and recover C₂H₂ from 50/50 or 1/99 C₂H₂/CO₂ and C₂H₂/C₂H₄ binary mixtures. *In situ* single crystal X-ray diffraction (SCXRD) combined with dispersion-corrected density functional theory (DFT-D) calculations reveals that the C₂H₂ binding site involves two *cis*-SiF₆²⁻ anions in close proximity (F...F distance of 7.16 Å), creating a new type of molecular trap that affords six uncoordinated fluoro moieties to chelate each C₂H₂ *via* sixfold C-H...F hydrogen bonds. This work therefore provides a new strategy for binding site engineering with selective C₂H₂ affinity to enable trace C₂H₂ capture.

Received 25th January 2025

Accepted 14th April 2025

DOI: 10.1039/d5sc00697j

rsc.li/chemical-science

Introduction

Acetylene (C₂H₂) is a high volume feedstock used for the production of various commodity chemicals.¹ The production of C₂H₂ involves partial combustion of methane or thermal cracking of hydrocarbons, inevitably generating impurities, for example, carbon dioxide (CO₂).² As a result, purification of C₂H₂ is a needed prerequisite for its downstream use.³ Capture of trace C₂H₂ impurities from gas mixtures is also of importance:

recovery of trace C₂H₂ is desirable in a circular economy and for safety, given that C₂H₂ is a flammable, explosive and toxic chemical;⁴ removal of trace C₂H₂ is a necessary procedure to produce polymer-grade ethylene (C₂H₄).⁵ Traditional purification technologies, such as organic solvent extraction (*e.g.*, *N,N*-dimethylformamide or acetone) or catalytic partial hydrogenation using noble metal catalysts,⁶ have high energy footprints and costs for waste solvent disposal. New approaches for trace C₂H₂ capture and separation are therefore desirable.⁷

Physisorption using porous materials is generally recognized as an efficient and energy-saving approach to C₂H₂ capture and separation.^{4,8} In this context, metal-organic materials such as metal-organic frameworks (MOFs)⁹ or coordination polymers (CPs)¹⁰ offer potential utility, thanks to their high surface area and modularity that allow for fine-tuning of the pore size, shape and chemistry.^{11–22} In this context, hybrid ultramicroporous materials (HUMs) composed of inorganic (*e.g.* SiF₆²⁻ – SIFSIX) and organic linker ligands have been reported to show benchmark trace C₂H₂ capture performance, thanks to their sub-nanometer pore size (<0.7 nm). In these HUMs, the C₂H₂ molecules are bound to the pore surface *via* strong hydrogen bonds because of the relatively protic nature of the CH moieties of acetylene. For example, the strong C-H...F hydrogen bonds

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† Electronic supplementary information (ESI) available. CCDC 2350654–2350657, 2361231–2361233. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d5sc00697j>

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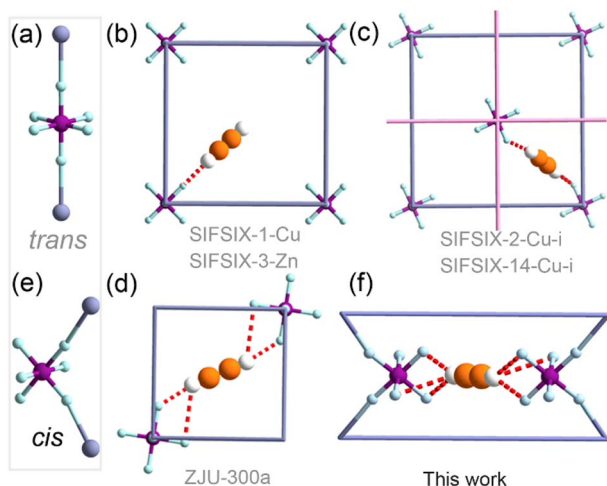


Fig. 1 SIFSIX anions can bridge metals in either *trans*- (a) or *cis*- (e) mode. C_2H_2 binding sites in *trans*-bridging materials can involve (b) a single $\text{C}-\text{H}\cdots\text{F}$ interaction, (c) dual $\text{C}-\text{H}\cdots\text{F}$ interactions or (d) quadruple $\text{C}-\text{H}\cdots\text{F}$ interactions. Herein, we report (f) a new type of binding site with sixfold $\text{C}-\text{H}\cdots\text{F}$ interactions involving three free fluoro atoms from each of the two *cis*-SIFSIX anions.

between C_2H_2 and the fluoro atom(s) of SiF_6^{2-} (SIFSIX) anions^{23–25} exemplify how molecular recognition can drive selectivity towards C_2H_2 .^{26–28}

In such HUM materials, SIFSIX anions invariably adopt a *trans*-bridging coordination mode (Fig. 1a, abbreviated hereafter as *trans*-SIFSIX).^{29–32} C_2H_2 binding in C_2H_2 selective *trans*-SIFSIX HUMs can be classified into one of the three modes (Fig. 1b–d).³³ a single $\text{C}-\text{H}\cdots\text{F}$ interaction (e.g., SIFSIX-1-Cu and SIFSIX-3);²³ dual $\text{C}-\text{H}\cdots\text{F}$ interactions from two SIFSIX anions (e.g. SIFSIX-2-Cu-i and SIFSIX-14-Cu-i);²⁴ quadruple $\text{C}-\text{H}\cdots\text{F}$ interactions from two SIFSIX anions involving bifurcated hydrogen bonds (e.g. ZJU-300a).³³ The *trans*-bridging coordination geometry of the SIFSIX anion in fact limits the free (uncoordinated) fluoro sites accessible for each C_2H_2 molecule (up to two per anion, as shown in Fig. 1d).

Recently, the first *cis*-bridging coordination mode of SIFSIX anions was observed in a flexible HUM (Fig. 1e).^{34–36} This *cis*-bridging exposes four identically oriented fluoro moieties to the pore walls²⁸ and promotes the possibility of formation of sixfold $\text{C}-\text{H}\cdots\text{F}$ interactions stemming from trifurcated hydrogen bonds (Fig. 1e and f). Herein, we report the first rigid HUM composed of *cis*-SIFSIX anions, $[\text{CuSiF}_6(\text{bidmb})_2]_n$ (SIFSIX-bidmb-Cu, where bidmb = 1,4-bis(1-imidazol-yl)-2,5-dimethyl benzene), and its sorption performance for C_2H_2 and related sorbates. As detailed herein, SIFSIX-bidmb-Cu is water stable with a binding site involving six fluoro moieties that offer outstanding separation performance for 50/50 or 1/99 mixtures of $\text{C}_2\text{H}_2/\text{CO}_2$ and $\text{C}_2\text{H}_2/\text{C}_2\text{H}_4$.

Results and discussion

Synthesis and crystal structures

Slow diffusion of a methanol solution of bidmb into a water solution of $\text{CuSiF}_6 \cdot 6\text{H}_2\text{O}$ after one month generated SIFSIX-

bidmb-Cu as purple block crystals suitable for single-crystal X-ray diffraction (SCXRD). Microcrystalline powders can be produced by a direct mixing method (Fig. S1†). The purity of bulk samples was confirmed by powder X-ray diffraction (PXRD) (Fig. S2†).

SCXRD analysis revealed that SIFSIX-bidmb-Cu crystallized in the triclinic space group $P\bar{1}$ (Table S1†). All Cu^{2+} cations adopt the same octahedral coordination geometry with four imidazolyl nitrogen atoms from four bidmb ligands at the equatorial positions and two fluorine atoms from SIFSIX at the axial positions (Fig. S3†). Cu^{2+} and bidmb generate the expected square lattice (**sql**) topology networks of formula $\text{Cu}(\text{bidmb})_2$ (Fig. 2a and b and S4†). These **sql** layers exhibit inclined interpenetration subtended by an angle of ca. 50° (Fig. 2c and d). SIFSIX anions crosslink the resulting networks through a *cis*-bridging mode to form a 3D self-catenated network with the simplified topological point symbol of $4^8.5^2.6^5$ (Fig. 2e and f), a topology that was first observed in SIFSIX materials (Fig. S5†). In SIFSIX-bidmb-Cu, the *cis*-SIFSIX anions and Cu centers generate an undulating CuSIFSIX chain (Fig. S6a†). Hydrogen bonds between fluorine atoms and hydrogen atoms of bidmb ($d_{\text{F}\cdots\text{H}} = 2.3\text{--}3.4 \text{ \AA}$) are present (Fig. S7 and S8†).

The structure contains one dimensional (1D) channels parallel to the crystallographic *b*-axis with a void space of 23.1% as determined by PLATON (Fig. 2g–i)³⁷ occupied by disordered MeOH and water molecules (Fig. S9†). The diameter of the 1D channel is ca. 4.6 \AA (Fig. S10†) and is inclined relative to the CuSiF_6 chain at an angle of 62° (Fig. S6b and c†). The connectivity and pore structure are therefore distinct from **pcu** topology SIFSIX materials in which 1D channels align parallel to MSIFSIX chains. The inclined structure herein results in SIFSIX anions exposed to the channel with $\text{F}\cdots\text{F}$ distances of 7.16 \AA (Fig. S11†). The electrostatic potential (ESP) of SIFSIX-bidmb-Cu and other benchmark SIFSIX materials was calculated to investigate the difference in pore chemistry arising from the coordination mode of SIFSIX anions (*trans* vs. *cis*). The ESP was mapped onto the Connolly surface with a probe radius of 1.7 \AA , and the same color gradation scale was used for comparison (Fig. 2j–n). Thanks to the strong electronegativity of fluorine atoms, the most negative electrostatic potential (ESP) is distributed in the region close to the SIFSIX anions in all SIFSIX materials. However, in SIFSIX-bidmb-Cu, the ESP is more negative than that of other SIFSIX materials.

In the first flexible *cis*-SIFSIX HUM, the **sql** layers are also composed of bis-imidazolyl ligands, but Cu centers are parallel to each other and pillared by *cis*-SIFSIX anions to form a **pcu** topology network.^{34–36} In SIFSIX-bidmb-Cu, **sql** layers are inclined to interpenetrate, reducing the pore size and changing the pore chemistry. HUMs formed by angular inorganic anions that pillar interpenetrated **sql** nets are unusual, but as exemplified by **mmo** nets containing tetrahedral anions (XO_4^{2-} , where $\text{X} = \text{W}$ and Mo)³⁸ that bridge three sets of mutually perpendicular **sql** nets (Fig. S5†), exceptional sorption performance is possible.

Characterization of SIFSIX-bidmb-Cu

TGA analysis revealed that SIFSIX-bidmb-Cu is thermally stable up to 230°C with guest molecules released by ca. 120°C (Fig. S12†).



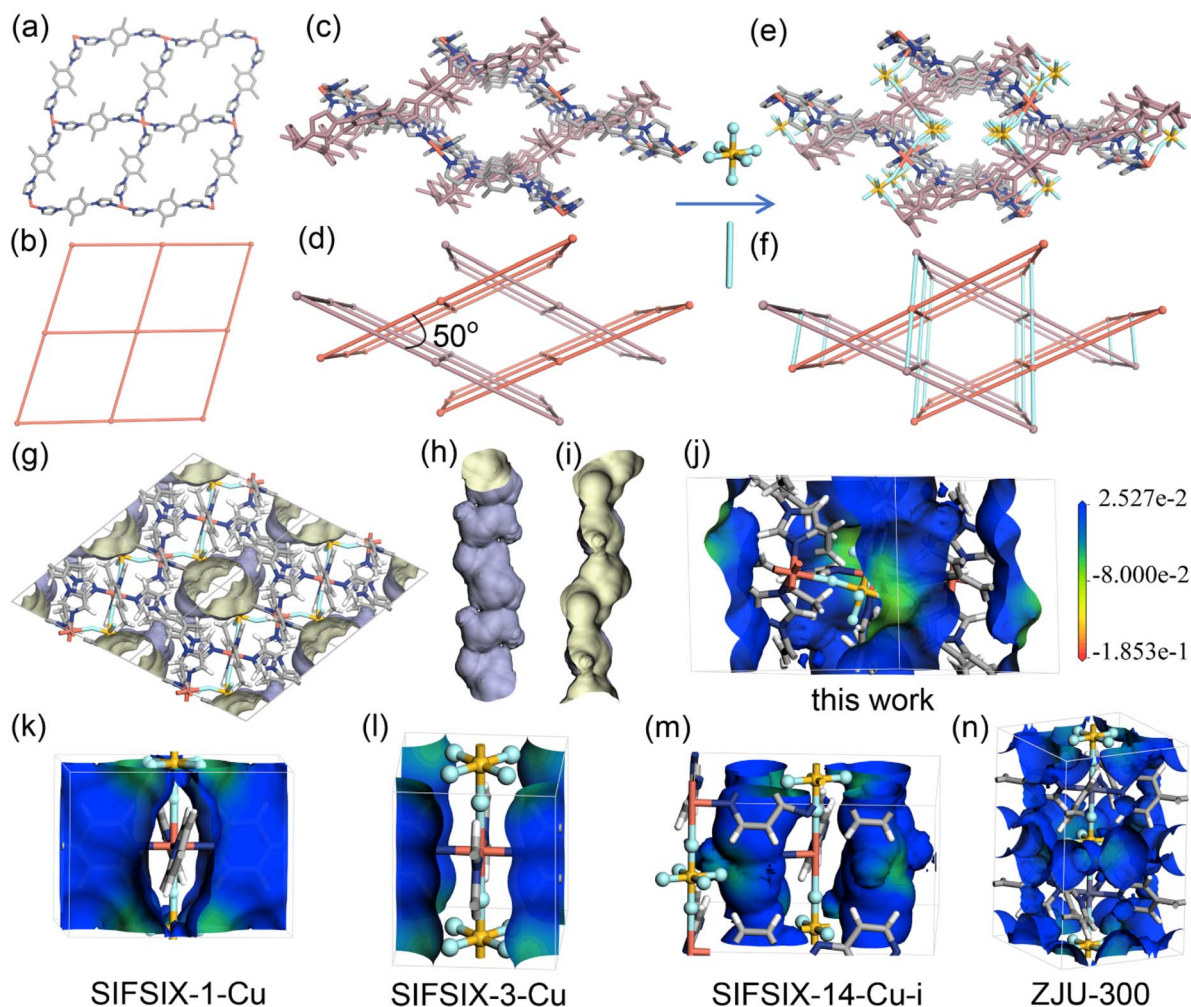


Fig. 2 (a–f) Structural and topological representation of the $\text{Cu}(\text{bidmb})_2$ sql layer (a and b), inclined interpenetrated sql nets (c and d) and the 3D self-catenated network formed by *cis*- SiF_6^{2-} anions pillaring interpenetrated sql nets (e and f). (g) The 1D channels along the *b*-axis, (h) the exterior, and (i) interior pore wall of a single 1D channel, represented by the Connolly surface with a probe radius of 1.7 Å. (j–n) Electrostatic potential (ESP) of SIFSIX-bidmb-Cu (j), SIFSIX-1-Cu (k), SIFSIX-3-Cu (l), SIFSIX-14-Cu-i (m) and ZJU-300 (n) mapped onto the Connolly surface. The gradation on the scale bar is in $\text{Ha}^* \text{electron}$ (–1).

Variable-temperature PXRD (VT-PXRD) data are consistent with SIFSIX-bidmb-Cu being a rigid sorbent (Fig. S13†). SCXRD analysis of the fully activated form, SIFSIX-bidmb-Cu', also suggests rigidity (Table S1†).

The permanent porosity of SIFSIX-bidmb-Cu' was tested by N_2 (77 K) and CO_2 (195 K) sorption experiments. Both gases exhibited type-I sorption isotherms with saturated uptake values of 5.79 and 4.51 mmol g^{-1} , respectively (Fig. 3a). Based on the 77 K N_2 adsorption isotherm, the Langmuir and Brunauer–Emmett–Teller surface areas were calculated to be 543.9 and 480 $\text{m}^2 \text{g}^{-1}$, respectively (Fig. S14 and S15†). The calculated pore volume, 0.16 $\text{cm}^3 \text{g}^{-1}$, is close to the theoretical value (0.18 $\text{cm}^3 \text{g}^{-1}$) on the basis of SCXRD. To our knowledge, SIFSIX-bidmb-Cu is the first rigid HUM (and only the second SIFSIX material) comprising *cis*-SIFSIX anions.^{34–36}

Single-component adsorption isotherms

The single-component sorption isotherms of SIFSIX-bidmb-Cu' for $\text{C}_2\text{H}_2/\text{CO}_2/\text{C}_2\text{H}_4$ were collected at 298 and 273 K. As shown in

Fig. 3b, each gas exhibits a type-I sorption isotherm at 298 K. The saturated uptake values of C_2H_2 , CO_2 and C_2H_4 at 1 bar were observed to be 2.65, 2.68, and 2.04 mmol g^{-1} , respectively. Notably, SIFSIX-bidmb-Cu showed much steeper C_2H_2 uptake than CO_2 and C_2H_4 at lower pressures (Fig. 3c). Substantially higher C_2H_2 uptake was exhibited at 0.01 bar (1.86 mmol g^{-1}) than for CO_2 (0.64 mmol g^{-1}) and C_2H_4 (0.33 mmol g^{-1}). The uptake of C_2H_2 at 0.01 bar, an indicator of the trace C_2H_2 capture ability of a sorbent, is higher than most SIFSIX materials (Fig. 3d and S16†), such as SIFSIX-1-Cu (0.45 mmol g^{-1}),²³ SIFSIX-3-Zn (0.85 mmol g^{-1}),²³ and SIFSIX-2-Cu-i (1.38 mmol g^{-1}),²³ comparable to the benchmark SIFSIX-14-Cu-i (1.83 mmol g^{-1}),²⁴ but is lower than ZJU-300a (3.23 mmol g^{-1}).³³ The PXRD patterns of SIFSIX-bidmb-Cu' collected at 273 and 298 K at 1 bar C_2H_2 are consistent with that calculated for SIFSIX-bidmb-Cu (Fig. S17†), further suggesting a rigid structure. Consecutive C_2H_2 adsorption–desorption experiments conducted at 298 K indicated that adsorbed C_2H_2 was not fully removed under vacuum at room temperature (about 50% uptake remained),



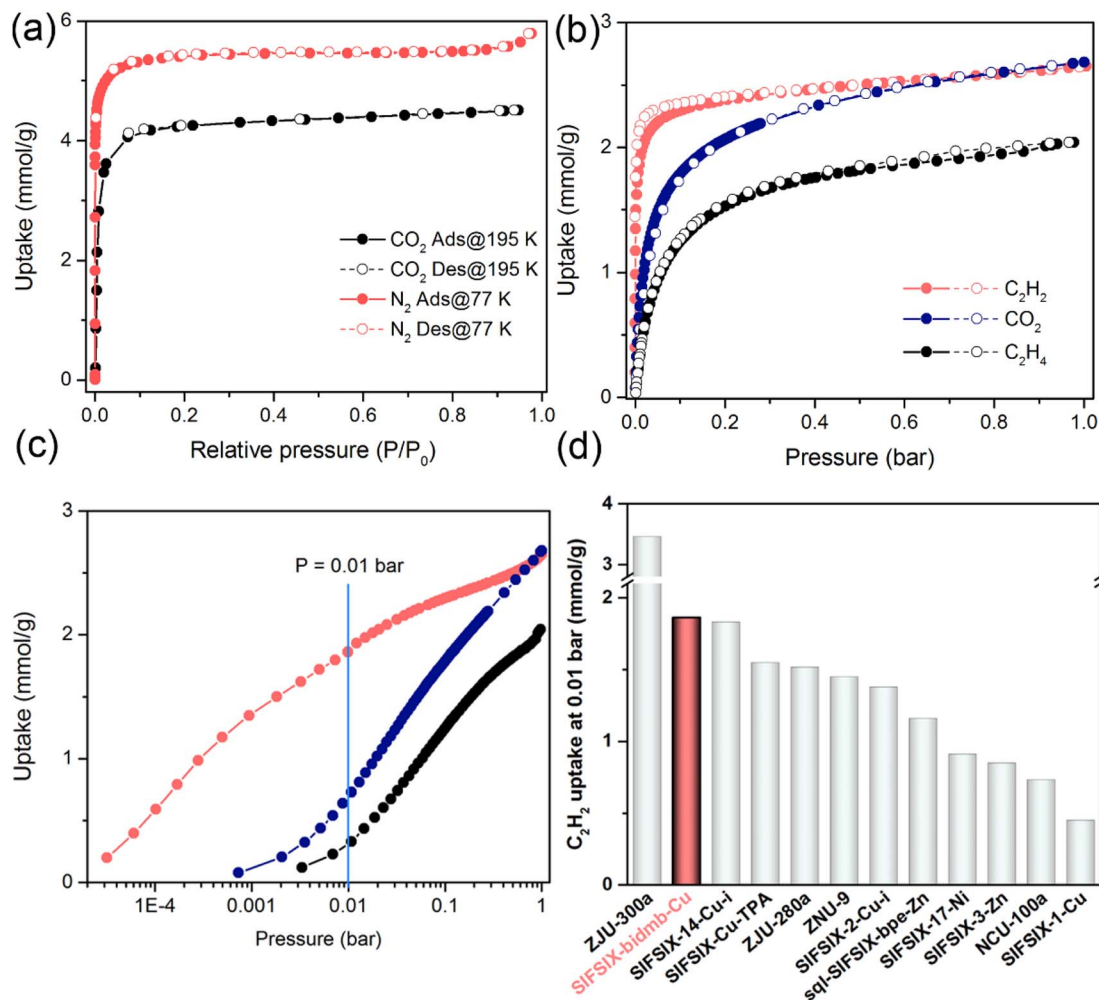


Fig. 3 (a) CO₂ (195 K) and N₂ (77 K) sorption isotherms. (b) and (c) C₂H₂, C₂H₄ and CO₂ sorption isotherms collected at 298 K. (d) Comparison of C₂H₂ uptake at 0.01 bar for SIFSIX-bidmb-Cu and other SIFSIX materials.

requiring heating to 120 °C (Fig. S18†). This provides evidence of the strong binding affinity for C₂H₂ and also suggests a method to capture and recover C₂H₂ *via* temperature control. Interestingly, the C₂H₂-saturated sample was regenerated under vacuum at room temperature after exposure to air for one year (Fig. S19–S21†). In contrast, multiple cycles of sorption experiments on CO₂ and C₂H₄ displayed negligible changes even if regeneration was implemented under high vacuum at room temperature (Fig. S22 and S23†), implying weaker interactions with SIFSIX-bidmb-Cu'. The time-dependent adsorption kinetics profiles of SIFSIX-bidmb-Cu' were also investigated at 298 K (Fig. S24†). SIFSIX-bidmb-Cu' exhibited a notably faster initial rate for C₂H₂ adsorption than CO₂, reaching 41.7% and 11.3% of 1 bar uptake in 40 seconds for C₂H₂ and CO₂, respectively. C₂H₂ adsorption reached equilibrium within 25 min, but CO₂ adsorption only reached *ca.* 40% of loading by this time.

The coverage-dependent isosteric heat of adsorption (Q_{st}) of SIFSIX-bidmb-Cu' for each gas derived from sorption isotherms indicated Q_{st} values at near-zero coverage as follows: C₂H₂ (55.7 kJ mol⁻¹) > C₂H₄ (38.1 kJ mol⁻¹) > CO₂ (20.1 kJ mol⁻¹)

(Fig. S25–S28†). Notably, the Q_{st} value of C₂H₂ at a low loading (55.7 kJ mol⁻¹) surpasses those of most prominent SIFSIX materials (Fig. S29†), including ZNU-9 (33.1 kJ mol⁻¹),³⁹ SIFSIX-21-Ni (37.9 kJ mol⁻¹),²⁹ SIFSIX-Cu-TPA (39 kJ mol⁻¹),⁴⁰ SIFSIX-14-Cu-i (40.0 kJ mol⁻¹),²⁴ SIFSIX-2-Cu-i (41.8 kJ mol⁻¹)²³ and SIFSIX-17-Ni (44.2 kJ mol⁻¹),⁴¹ and is comparable to those of UTSA-300a (57.6 kJ mol⁻¹),³⁰ NCU-100a (60.5 kJ mol⁻¹)⁴² and ZJU-300a (61.1 kJ mol⁻¹).³³ Conversely, the $Q_{st}(\text{CO}_2)$ value at a low loading (20.1 kJ mol⁻¹) is low, being marginally higher than that of SIFSIX-21-Ni (19.8 kJ mol⁻¹).²⁹ Consequently, the difference in Q_{st} values between C₂H₂ and CO₂ ($\Delta Q_{st, \text{C}_2\text{H}_2-\text{CO}_2}$), 35.6 kJ mol⁻¹, is to our knowledge the highest reported value for SIFSIX materials (SIFSIX-2-Cu-i = 9.9 kJ mol⁻¹,⁴³ SIFSIX-22-Zn = 11.5 kJ mol⁻¹,⁴⁴ SIFSIX-Cu-TPA = 13.4 kJ mol⁻¹,⁴⁰ and SIFSIX-21-Ni = 18.1 kJ mol⁻¹) (Fig. S30 and S31†).²⁹ Furthermore, this ΔQ_{st} value is higher than the values of other top-performing sorbents like SOFOUR-1-Zn (24 kJ mol⁻¹),⁴⁴ SOFOUR-TEPE-Zn (26.3 kJ mol⁻¹),⁴⁵ Cu(bpy)NP (26.2 kJ mol⁻¹)⁴⁶ and SNNU-98-Mn (31 kJ mol⁻¹),⁴⁷ comparable to that of Ni(4-DPDS)₂CrO₄ (38.4 kJ mol⁻¹),⁴⁸ but inferior to those of sorbents with open metal centres such as ATC-Cu (43.6 kJ mol⁻¹)⁴⁹ and Cu^I@UiO-

66-(COOH)₂ (45.6 kJ mol⁻¹).⁵⁰ The difference between $Q_{\text{st}}(\text{C}_2\text{H}_2)$ and $Q_{\text{st}}(\text{C}_2\text{H}_4)$, $\Delta Q_{\text{st},\text{C}_2\text{H}_2-\text{C}_2\text{H}_4}$, of 17.6 kJ mol⁻¹ is also comparable to that of the benchmark SIFSIX material ZJU-300a (21.1 kJ mol⁻¹)³³ and superior to those of SIFSIX-3-Ni (10.2 kJ mol⁻¹),⁵¹ SIFSIX-2-Cu-i (11.2 kJ mol⁻¹)⁴³ and SIFSIX-14-Cu-i (13 kJ mol⁻¹) (Fig. S32†).²⁴ These Q_{st} and ΔQ_{st} values are indicative of the relative binding strength of SIFSIX-bidmb-Cu' for these sorbates.

Separation properties

To explore the separation properties of SIFSIX-bidmb-Cu', selectivities for C₂H₂/CO₂ and C₂H₂/C₂H₄ gas mixtures in different compositions at 100 kPa and 298 K were calculated from the ideal adsorbed solution theory (IAST) (Fig. S33–S42†). For both C₂H₂/CO₂ and C₂H₂/C₂H₄, the selectivity values indicate potential for trace removal of C₂H₂. For 50/50 (v/v) C₂H₂/CO₂ and 1/99 C₂H₂/C₂H₄, the two gas ratios most commonly investigated, the selectivity values were determined to be 20.3 and 140.2, respectively (Fig. S36†). The equimolar C₂H₂/CO₂ selectivity of SIFSIX-bidmb-Cu' is higher than most SIFSIX materials (Fig. S43†), e.g. SIFSIX-Cu-TPA (5.3),⁴⁰ SIFSIX-21-Ni (10),²⁹ ZNU-9 (10.3),³⁹ SIFSIX-17-Ni (11.7)⁴¹ and ZJU-280a (18.1),⁵² and is only lower than those sorbents that exhibit molecular sieving, including UTSA-300a (743)³⁰ and NCU-100a (1786.6).⁴² The 1/99 C₂H₂/C₂H₄ selectivity of SIFSIX-bidmb-Cu' is also superior to that of most SIFSIX materials (Fig. S44†), including SIFSIX-3-Ni (5.03),⁵¹ SIFSIX-2-Cu (6),²³ SIFSIX-1-Cu (10.63),²³ ZNU-9 (11.64),³⁹ SIFSIX-2-Cu-i (44.5),⁴³ ZJU-280a (44.5),⁵² and sql-SIFSIX-bpe-Zn (53.1),⁵³ but inferior to that of the molecular sieve SIFSIX-14-Cu-i (6320).²⁴ Notably, SIFSIX-bidmb-Cu' ranks second among sorbents that simultaneously exhibit 50/50 C₂H₂/CO₂ and 1/99 C₂H₂/C₂H₄ selectivities, only being inferior to SIFSIX-TEPE-Cu (Fig. S45†).⁵⁴

Since 2016, when SIFSIX HUMs were found to offer superior acetylene sorption/separation, different strategies have been tried to improve the sorption/separation performance, including pore size control *via* interpenetration (e.g. SIFSIX-2-Cu *vs.* SIFSIX-2-Cu-i),^{23,24} organic linker functionalization⁴¹ and anion replacement to fine-tune pore chemistry (e.g. SiF₆²⁻ to TiF₆²⁻).⁵¹ In contrast, C₂H₂ capture performance improvement *via trans-to-cis* bridging coordination mode change of SIFSIX anions has not been explored. This research work fills the gap and shows that *cis*-SIFSIX anions can also be favorable for the construction of C₂H₂-selective sorbents with *cis*-SIFSIX anions, offering more fluoro binding sites than *trans* analogs. The selectivity for 1/99 C₂H₂/C₂H₄ has been improved from 44.5 (SIFSIX-2-Cu-i) to 140.2 (SIFSIX-bidmb-Cu), while the selectivity for 50/50 C₂H₂/CO₂ has been increased to 20.3 (SIFSIX-bidmb-Cu) from 6.5 (TIFSIX-2-Cu-i). The *cis*-bridging mode of SIFSIX anions resembles that of the tetrahedral anions XO₄²⁻ (X = W, Mo, S *etc.*), which are increasingly used in HUMs.^{38,44,45}

Dynamic breakthrough experiments

The experimental separation performance of SIFSIX-bidmb-Cu' for C₂H₂/CO₂ and C₂H₂/C₂H₄ mixtures was tested by transient column breakthrough experiments performed at 298 K and 100

kPa (Fig. 4a–d). SIFSIX-bidmb-Cu' showed excellent C₂H₂ separation performance for both 50/50 and 1/99 binary mixtures. For C₂H₂/CO₂ mixtures, CO₂ breakthrough occurred first with retention times of 59.7 and 128.1 min g⁻¹ for the 50/50 and 1/99 mixtures, respectively (Fig. 4a and b). In contrast, the corresponding C₂H₂ breakthrough occurred after retention times of 98.6 and 1591.2 min g⁻¹, respectively. The breakthrough times for C₂H₂ and CO₂ were 38.9 and 1463.1 min g⁻¹ for 50/50 and 1/99 gas mixtures, respectively. The calculated C₂H₂ uptake values were 2.48 and 1.38 mmol g⁻¹ for 50/50 and 1/99 mixtures, respectively. The former is comparable to that of the benchmark materials, including JCM-1 (2.2 mmol g⁻¹),⁵⁵ NKMOF-1-Ni (2.48 mmol g⁻¹),⁵⁶ Cu^I@UiO-66-(COOH)₂ (2.89 mmol g⁻¹)⁵⁰ and Ni(4-DPDS)₂CrO₄ (2.96 mmol g⁻¹).⁴⁸ Captured C₂H₂ can be recovered with high purity (>99% purity) through heating the saturated sample at 120 °C after removal of co-adsorbed CO₂ *via* a period of helium flush at room temperature (the longer the flushing time, the purer the concentration of C₂H₂) (Fig. 4e and f). The temperature-controlled recovery of C₂H₂ is in good agreement with the sorption experiments where heating at 120 °C was needed to fully remove the strongly adsorbed C₂H₂ molecules. To our knowledge, SIFSIX-bidmb-Cu is just the second sorbent that yields high purity C₂H₂ (>99% purity) from a 1/99 C₂H₂/CO₂ stream on desorption,⁴⁴ which is relevant for recycling of C₂H₂, safety and environmental protection. The equimolar C₂H₂/CO₂ separation factor for SIFSIX-bidmb-Cu, a criterion for evaluating the separation potential of sorbents, was calculated to be 7.8, higher than those of SIFSIX-Cu-TPA (1.97),⁴⁰ FJUT-1 (5.17),⁵⁷ Cu^I@UiO-66-(COOH)₂ (3.4),⁵⁰ JCM-1 (4.4),⁵⁵ FeNi-M'MOF (1.7),⁵⁸ and Ni(4-DPDS)₂CrO₄ (6.7).⁴⁸

For 50/50 and 1/99 C₂H₂/C₂H₄ binary mixtures, the C₂H₄ gas broke through the adsorption bed with high-purity at 39.7 and 41.5 min g⁻¹, while the breakthrough time of C₂H₂ was longer at 106.2 and 1452.4 min g⁻¹, corresponding to breakthrough intervals of 66.5 and 1410.9 min g⁻¹, respectively (Fig. 4c and d). The calculated C₂H₂ capture capacities were 2.68 and 1.22 mmol g⁻¹ for 50/50 and 1/99 mixtures, respectively. The productivity of 99.999% pure C₂H₄ was calculated to be 84.6 mmol g⁻¹, which is comparable to the benchmark SIFSIX-14-Cu-i (84.6 mmol g⁻¹) and higher than those of SIFSIX-3-Zn (1.94 mmol g⁻¹), MUF-17 (8.57 mmol g⁻¹),⁵⁹ SIFSIX-1-Cu (12.4 mmol g⁻¹),²³ SIFSIX-dps-Cu (14.9 mmol g⁻¹),⁴³ ZNU-9 (48.57 mmol g⁻¹)³⁹ and Cu(bpy)NP (20.57 mmol g⁻¹).⁴⁶ According to the breakthrough curves, the C₂H₂/C₂H₄ dynamic selectivity (1/99) of SIFSIX-bidmb-Cu was calculated to be 105.7 (Fig. 4h), notably higher than previously studied sorbents, including SIFSIX-3-Zn (9), SIFSIX-1-Cu (11), and SIFSIX-2-Cu-i (45),²³ comparable to Ni@FAU (97)⁸ and SIFSIX-14-Cu-i (91),²⁴ but lower than ZJU-300a (264).³³ Similarly, captured C₂H₂ can be recovered with high purity (>99% purity) *via* heating the sample at 120 °C after a helium flush at room temperature (Fig. 4g), thus providing a facile approach to obtain pure C₂H₂ and C₂H₄. Additionally, the breakthrough performance for the 10/90 C₂H₂/C₂H₄ mixture was investigated at 298 K and 100 kPa with a total gas flow rate of 2 mL min⁻¹ (Fig. 4i), resulting in retention times of 15.8 and 216.3 min g⁻¹ for C₂H₄ and C₂H₂, respectively, and



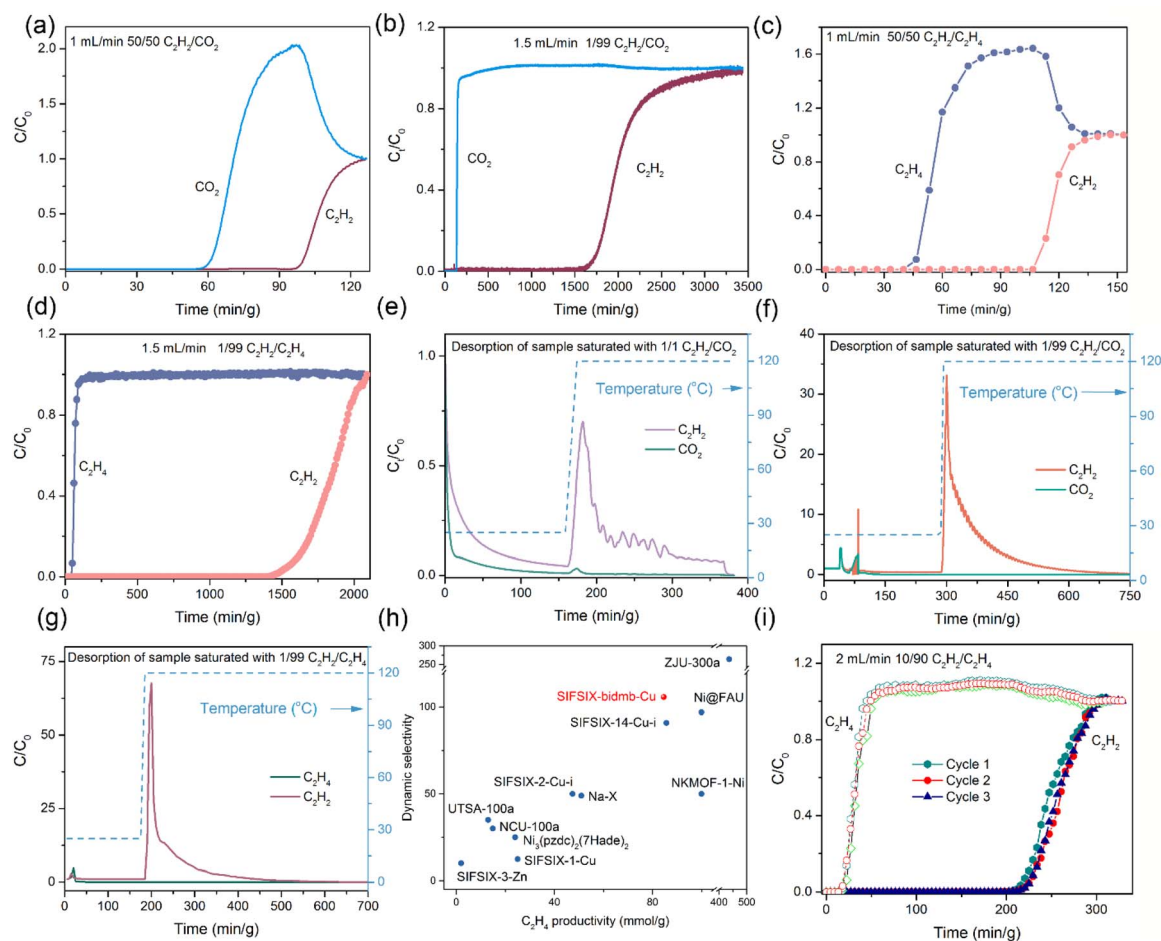


Fig. 4 (a–d) Experimental breakthrough curves of equimolar C_2H_2/CO_2 (a), 1/99 C_2H_2/CO_2 (b), equimolar C_2H_2/C_2H_4 (c) and 1/99 C_2H_2/C_2H_4 (d) mixtures on SIFSIX-bidmb-Cu'. (e) and (f) Desorption curves of samples saturated with 1/1 (e) and 1/99 C_2H_2/CO_2 (f) and 1/99 C_2H_2/C_2H_4 (g) mixtures. (h) Comparison of C_2H_4 productivity versus dynamic selectivity (1/99 C_2H_2/C_2H_4) for SIFSIX-bidmb-Cu' and other benchmark sorbents. (i) Three consecutive cycles of breakthrough curves for a 10/90 C_2H_2/C_2H_4 binary mixture on SIFSIX-bidmb-Cu'.

a breakthrough time lag of 200.5 min g^{-1} . Consecutive breakthrough experiments for 10/90 C_2H_2/C_2H_4 on SIFSIX-bidmb-Cu' revealed no performance loss, suggesting recyclability which was supported by PXRD data (Fig. S46†). The breakthrough experiments for 50/50 C_2H_2/CO_2 and 1/99 C_2H_2/C_2H_4 mixtures conducted at higher gas flow rates (5 mL min^{-1}) were in good agreement with those at lower flow rates, demonstrating that the separation performance was unaffected by gas flow rates and highlighting the fast sorption kinetics of C_2H_2 (Fig. S47†).

Binding site identification via SCXRD and DFT-D calculations

In order to gain insight into the difference in binding affinity between the three sorbates and the framework, the binding sites of SIFSIX-bidmb-Cu' were explored via SCXRD on gas-loaded single crystals in combination with dispersion-corrected density functional theory (DFT-D) calculations (Fig. 5). *In situ* SCXRD studies of SIFSIX-bidmb-Cu' under 1 bar gas pressure at 298 K using an environmental gas cell were conducted first. In all cases, the gas molecules could not be modeled crystallographically (most likely owing to their high

thermal motion at 298 K); however, the most probable locations of the guest molecules within the host framework can be inferred from difference electron density maps (Fig. S48–S52†). In order to further visualize the binding sites, SCXRD studies at 100 K were performed on crystals loaded with C_2H_2 and CO_2 at 298 K. $C_2H_2@SIFSIX-bidmb-Cu'$ was found to exhibit two types of binding sites (site I and site II) for C_2H_2 molecules, each located between two adjacent SiF_6^{2-} anions (Fig. 5a–d and g). The C_2H_2 molecule at each site is chelated by two SiF_6^{2-} anions in an end-on fashion through $C\equiv C-H\cdots F$ hydrogen bonding ($d_{C-H\cdots F}$: 2.16–3.16 Å), and there are van der Waals (vdW) interactions between C_2H_2 and bidmb ligands ($d_{C-H\cdots C\equiv C}$: 3.19–3.63 Å) (Fig. 5d and g), affording a zig-zag $SiF_6^{2-}\cdots C_2H_2(I)\cdots SiF_6^{2-}\cdots C_2H_2(II)\cdots SiF_6^{2-}$ chain along the channel direction (*b*-axis) (Fig. 5a–c). At site I, two parallel and centrosymmetric C_2H_2 molecules are simultaneously chelated by two SIFSIX anions, wherein each is unsymmetrically bound to four terminal fluorine atoms from two SiF_6^{2-} anions (Fig. 5d and g). At site II, the C_2H_2 molecule is located at an inversion center and is symmetrically bound to six terminal fluorine atoms from two SIFSIX anions (Fig. S48†).⁶⁰ For the two binding sites, the DFT-D

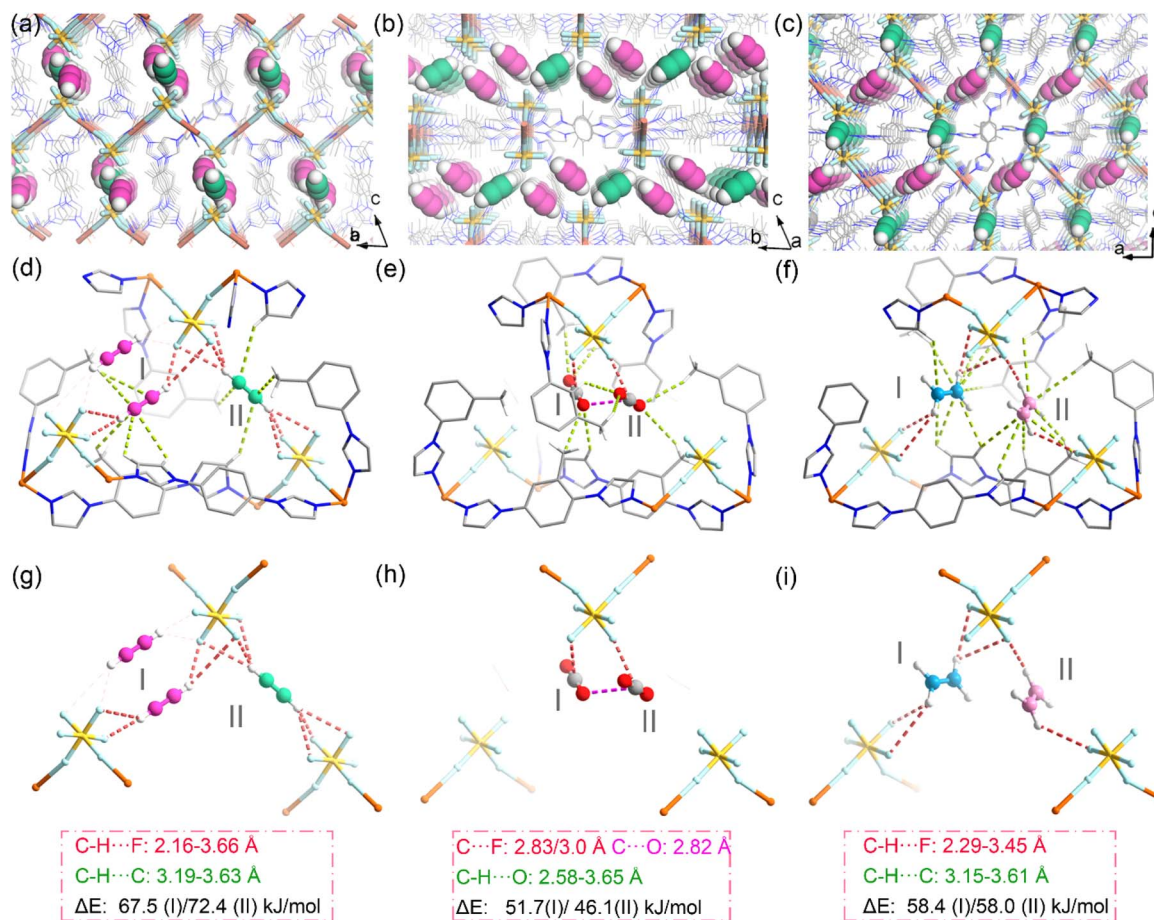


Fig. 5 (a–c) The packing of adsorbed C₂H₂ molecules in SIFSIX-bidmb-Cu' viewed in different directions based on SCXRD results. (d–f) The binding sites and configurations of C₂H₂ (d), CO₂ (e) and C₂H₄ (f). (g–i) The interactions between SiF₆²⁻ anions and C₂H₂ (g), CO₂ (h) and C₂H₄ (i).

calculated static binding energies (ΔE) are ~ 67.5 and ~ 72.4 kJ mol⁻¹, respectively, which are higher than the strongest calculated C₂H₂ binding strength in SIFSIX materials (~ 52.9 kJ mol⁻¹ in SIFSIX-2-Cu-i,⁴³ ~ 56.0 kJ mol⁻¹ in SIFSIX-14-Cu-i²⁴ and $\sim 51.7/56.8$ kJ mol⁻¹ in UTSA-300a³⁰).

For CO₂@SIFSIX-bidmb-Cu', the CO₂ molecules could not be modeled adequately, consistent with the low Q_{st} of CO₂ adsorption. The difference electron density maps obtained from SCXRD studies at 298 K and 100 K on CO₂-loaded crystals reveal similar guest distribution regions (Fig. S50†). We performed DFT-D calculations to determine the binding sites of CO₂ and C₂H₄. Three CO₂ binding sites (site I, site II and site III) were identified (Fig. 5e and h). CO₂ molecules at sites I and II lie in similar positions but with different orientations than that of C₂H₂. Each CO₂ molecule is bound to one (site II and site III) or two (site I) terminal fluorine atoms from a SiF₆²⁻ anion in a side-on mode through O=C...F electrostatic interactions ($d_{O=C...F}$: 2.86–3.18 Å), accompanied by vdW interactions between CO₂ and bidmb ligands ($d_{C-H...O=C}$: 2.50–3.70 Å). The CO₂ molecules on site I and site II should be half-occupied, as they are too close to the inversion centers of the structure (Fig. S53†). The DFT-D calculated static binding energies (ΔE) for the three CO₂ binding sites were determined to be ~ 48.0 ,

~ 45.1 and ~ 52.8 kJ mol⁻¹, respectively. We attribute this to the lack of synergetic fluorine–CO₂ binding interactions.

For C₂H₄, two binding sites were identified (site I and site II), displaying similar positions and orientations to those of C₂H₂ (Fig. 5f and i). On each site, C₂H₄ is chelated by SIFSIX pairs in an end-on mode through =C-H...F hydrogen bonding ($d_{=C-H...F}$: 2.29–3.45 Å), which is accompanied by vdW interactions between C₂H₄ and bidmb ligands ($d_{C-H...C=C}$: 3.15–3.61 Å). Each SiF₆²⁻ anion at site I and site II provides two and one free fluoro atoms that bind to one hydrogen atom of C₂H₄, respectively. Similarly, the C₂H₄ molecule at site II should be half-occupied, as it resides close to the inversion center (Fig. S54†). For the two C₂H₄ binding sites, the DFT-D calculated static binding energy (ΔE) was determined to be ~ 58.0 and ~ 58.4 kJ mol⁻¹, respectively. We attribute these lower energies to weaker C-H...F hydrogen bonds (C₂H₂ is more acidic than C₂H₄). The C₂H₂ binding sites from SCXRD data collected at 100 K and those of CO₂ and C₂H₄ calculated from DFT-D are in good agreement with the guest locations inferred from the difference electron density maps of *in situ* SCXRD at 298 K (Fig. S49–S52†). The DFT-D calculated static binding energy (ΔE) is in the sequence of C₂H₂ > C₂H₄ > CO₂ and matches well with the order of Q_{st} values from sorption isotherms.



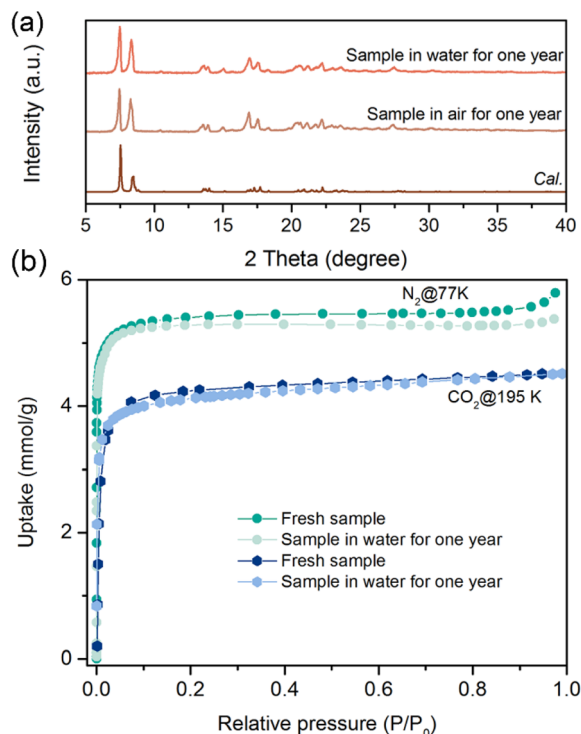


Fig. 6 (a) Comparison of the calculated PXRD pattern with those of samples treated in air or water for one year. (b) Comparison of CO₂ (195 K) and N₂ (77 K) isotherms of pristine sample with those of samples treated in air or water for one year.

In *trans*-SIFSIX HUMs, due to the symmetric coordination mode of the SiF₆^{2−} anion, the four free (uncoordinated) F sites are equally distributed around the coordination axis. The free rotation of *trans*-SIFSIX anions around the coordination axis will modulate the number of free F sites located on the pore surface at a binding site. For example, in SIFSIX-1-Cu,²³ only one free F site orients toward the channel, and thus a single C–H...F hydrogen bonding was observed (the pore size is large too). When the pore size was reduced by, for instance, interpenetration, the distance between two anions is less (e.g., SIFSIX-2-Cu-i and SIFSIX-14-Cu-i),^{23,24} and dual C–H...F hydrogen bonds can occur. When the rotation of *trans* anions results in two free F sites located on the pore surface, each anion can provide two free F sites for one C₂H₂, and totally four C–H...F hydrogen bonds can be formed (e.g. ZJU-300a).³³ Due to geometry limitations, at most, two of the four F sites of a *trans*-SiF₆^{2−} anion can be accessible for one C₂H₂ molecule. Symmetric bridging of *trans*-SIFSIX anions can result in strong CO₂ trapping (e.g. SIFSIX-3-Ni).¹⁶ However, for *cis*-SIFSIX anions, the four free F sites are located on the same side of the coordination axis and one C₂H₂ molecule can simultaneously engage with three F sites of each anion. We attribute this feature to the highly selective binding of C₂H₂ vs. both CO₂ and C₂H₄ observed herein.

Water stability tests

SIFSIX materials sustained by *trans*-SIFSIX anions (e.g. SIFSIX-14-Cu-i) have been reported to undergo degradation or phase

transformation when exposed to humidity.⁶¹ Conversely, SIFSIX-bidmb-Cu bulk samples of SIFSIX-bidmb-Cu retained crystallinity after one year of exposure to both liquid water and ambient humidity, as confirmed by PXRD (Fig. 6a) and VT-PXRD (Fig. S55†) experiments. The adsorption isotherms of CO₂ (195 K), N₂ (77 K), C₂H₂ (298 K), CO₂ (298 K) and C₂H₄ (298 K) of samples exposed to air or water revealed negligible differences vs. pristine materials (Fig. 6b, S56 and S57†). Water stability was also demonstrated by consecutive water adsorption–desorption experiments that revealed no uptake loss even after ten cycles and no structural alteration according to PXRD (Fig. S58 and S59†). The high hydrolytic stability of SIFSIX-bidmb-Cu is consistent with that of the first *cis*-SIFSIX sorbent.³⁴ The use of imidazole linkers could also be a factor⁶² along with self-catenation.⁶³

Conclusions

This work reveals that *cis*-bridging SIFSIX anions can offer different sorbate binding sites relative to *trans* variants, in this case with high C₂H₂ affinity and low CO₂ affinity. In addition, this is the first example of a *cis*-bridging SIFSIX rigid HUM, and it exhibits C₂H₂ molecular traps enabled by multiple fluoro binding sites that exhibit relatively weak CO₂ and C₂H₄ binding. Overall, the combination of high C₂H₂ uptake at 0.01 bar, high *Q*_{st} for C₂H₂, large difference in *Q*_{st} between C₂H₂ and CO₂, and high C₂H₂/CO₂ and C₂H₂/C₂H₄ selectivities make SIFSIX-bidmb-Cu a leading sorbent for trace C₂H₂ capture. This work opens up a new avenue for the design and construction of SIFSIX materials with distinct structural features and pore chemistry. Other fluorinated linkers will be targeted in order to fine-tune C₂H₂ binding affinity.

Data availability

The data supporting this article have been included as part of the ESI.† The crystallographic information can be found in the ESI† associated with this work and at the Cambridge Crystallographic Data Center under deposition numbers 2350654–2350657, 2361231–2361233, via https://www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

Author contributions

B. Q.-Song and M. Y. Gao contributed equally to this work and carried out the synthesis, characterization analysis, sorption and separation measurements, and writing – original draft; C.-H. Deng and Q.-Y. Yang assisted with the gas sorption measurement and data analysis; S.-Q. Wang and S. Darwish assisted with the characterization analysis; L. M. van Wyk, A. C. Eaby and L. J. Barbour conducted the *in situ* single crystal X-ray diffraction measurements; D. Li and S.-J. Qin performed computational simulations; B.-Q. Song, Y. -L. Peng and M. J. Zaworotko carried out the methodology, supervision, and



writing – review & editing; all authors contributed to preparing the manuscript.

Conflicts of interest

The authors declare no competing financial interests.

Acknowledgements

The authors acknowledge the support of National Natural Science Foundation of China (Nos. 22201025 and 22371221), start-up funding from Chengdu University of Technology (10912 KYQD2022-09754) and the European Research Council (ADG 885695).

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