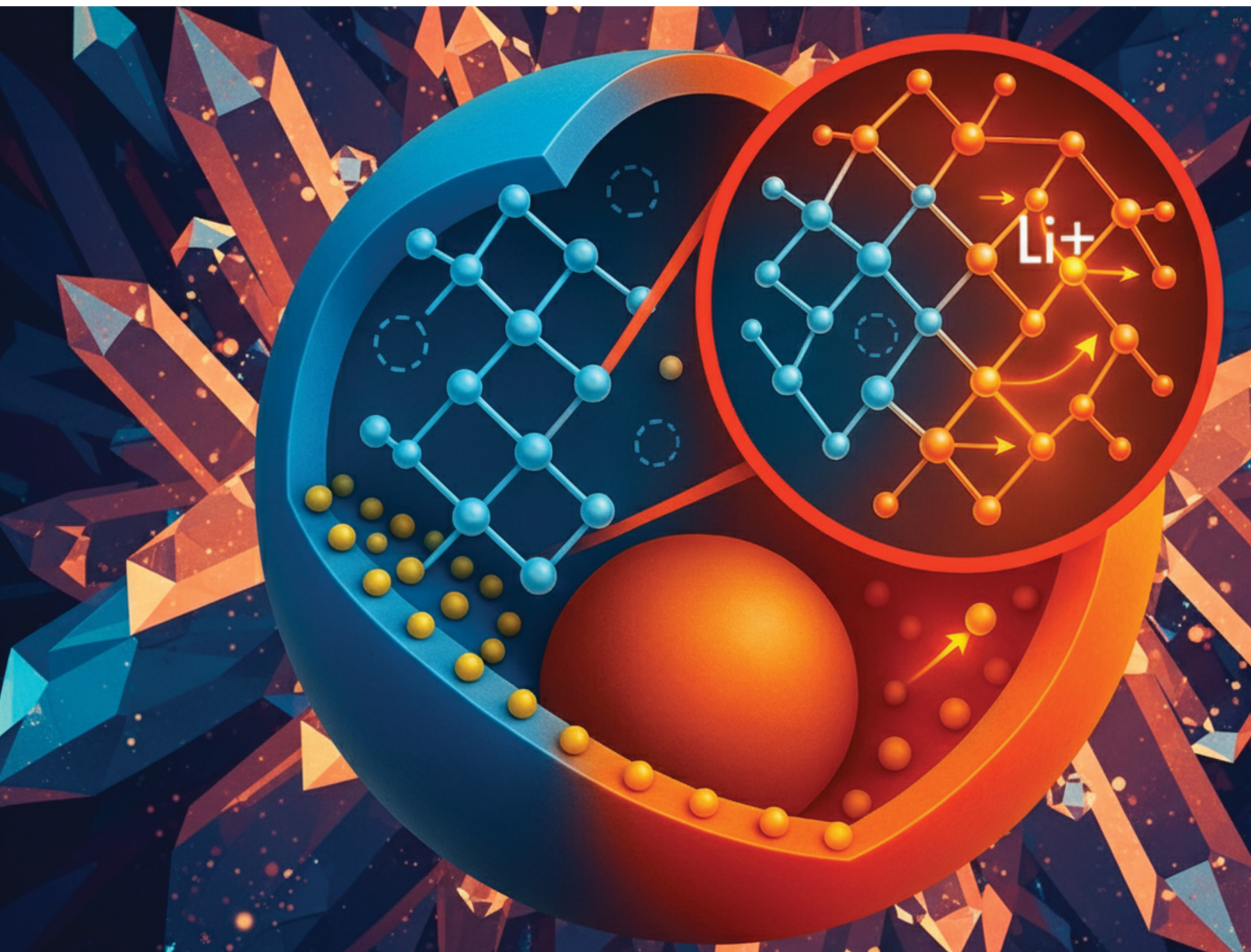


Materials Advances

Volume 6
Number 19
7 October 2025
Pages 6599–7128

rsc.li/materials-advances



ISSN 2633-5409

PAPER

Ha M. Nguyen, Carlos Wexler *et al.*
Vacancy-induced modification of the electronic band
structure of LiBO_2 materials as cathode surface coatings
for lithium-ion batteries

Cite this: *Mater. Adv.*, 2025, 6, 6682

Vacancy-induced modification of the electronic band structure of LiBO_2 materials as cathode surface coatings for lithium-ion batteries

Ha M. Nguyen,^{†*ab} Carson D. Ziemke,^{†ac} Narendrakumar Narayanan,^b Sebastian Amaya-Roncanci,^{id e} John Gahl,^b Yangchuan Xing,^{id ad} Thomas W. Heitmann^{abc} and Carlos Wexler^{id *ac}

LiBO_2 is an electronic insulator and a promising multi-functional material, notably as an effective surface coating for stabilizing high-voltage cathodes in lithium-ion batteries. Despite its potential, the underlying mechanisms of lithium-ion and electron transports in the presence of lattice vacancies remain inadequately understood. This is particularly important as such lattice vacancies may either facilitate or impede charge transport. A detailed understanding of these mechanisms is essential for the rational design and optimization of LiBO_2 -based materials. In our previous work [C. Ziemke, H. M. Nguyen, S. Amaya-Roncanci, J. Gahl, Y. Xing, T. W. Heitmann and C. Wexler, Formation of Lattice Vacancies and Their Effects on Li-ion Transport in LiBO_2 Crystals: A Comparative *ab initio* Study, *J. Mater. Chem. A*, 2025, **13**, 3146–3162], we used density functional theory (DFT) calculations to investigate the impact of lattice vacancies on Li-ion transport in both tetragonal (t-LBO) and monoclinic (m-LBO) polymorphs of LiBO_2 , revealing that B vacancies in either polymorph enhanced lithium-ion transport. In contrast, in this study we used DFT calculations to examine the effects of lattice vacancies on the electronic properties of both t-LBO and m-LBO polymorphs, focusing on the electronic band structure. Our analysis shows that B vacancies can enhance the electronic insulation of t-LBO while improving the electronic conduction of m-LBO. This finding reveals a strongly polymorph-dependent material selection, specifically the generation of B vacancies in LiBO_2 may enable t-LBO to function as a promising solid electrolyte (i.e., both a good ionic conductor and a good electronic insulator) and enhance the performance of m-LBO as a conformal cathode coating (i.e., both a good ionic conductor and a good electronic conductor) in lithium-ion batteries. Therefore, generating B vacancies, such as by neutron irradiation, would offer a viable strategy to improve the functionality of LiBO_2 as a promising material for energy storage applications. Overall, while focusing mainly on electron transport, this study complements our earlier work on ion transport and establishes a foundational understanding for future investigations of crucial aspects of LiBO_2 , such as electrochemical stability, mechanical robustness, and interfacial behavior.

Received 8th May 2025,
Accepted 18th July 2025

DOI: 10.1039/d5ma00458f

rsc.li/materials-advances

1. Introduction

Lithium metaborate (LiBO_2) is an electronic insulator that has drawn increasing research attention due to its multiple technological applications.^{1–4} These include the surface coating of the cathode of liquid electrolyte-based lithium-ion batteries (LIBs).^{5–8} The performance and efficiency of LIBs are critically dependent on the properties of the materials used in their construction, particularly those of the cathode and its coating.^{9–12} High-voltage cathode materials, which have been developed for LIBs operating at voltages as high as 3 to 6 V, include layered oxides LiMO_2 ($\text{M} = \text{Ni}_x\text{Mn}_{1-x-y}\text{Co}_y$ with $0 \leq x, y \leq 1$), spinel oxides LiM_2O_4 ($\text{M} = \text{Ni}_x\text{Mn}_{2-x}$, $0 \leq x \leq 2$), and

^a Materials Science and Engineering Institute, University of Missouri, Columbia, MO 65201, USA. E-mail: hn4gg@missouri.edu, wexlerc@missouri.edu^b University of Missouri Research Reactor (MURR), University of Missouri, Columbia, MO 65203, USA^c Department of Physics and Astronomy, University of Missouri, Columbia, MO 65201, USA^d Department of Chemical and Biomedical Engineering, University of Missouri, Columbia, MO 65201, USA^e Escuela de Física, Universidad Pedagógica y Tecnológica de Colombia, Tunja, Colombia[†] Authors with equal contributions.

polyanion compounds like LiMXO_4 ($M = \text{Fe, Ni, Co, and Mn; } X = \text{B, P, Si, W, Mo, etc.}$).^{11,13} However, such a high-voltage window can cause serious stability problems for the liquid electrolyte, the cathode itself, and hence for the overall performance of the LIB.^{11,13,14} For example, metal dendrite formation on the cathode surface, cathode surface degradation, decomposition of unstable components of the liquid electrolyte, and side reactions between the cathode and liquid electrolyte are some of the problems.^{11,13–15}

These problems have challenged the research community to seek effective strategies to mitigate them. To date, two of the most-studied approaches for the materials chemistry of LIBs are: (i) modification of the liquid electrolyte by means of additives in order to improve the stability of the solid electrolyte interphase (SEI) between the liquid electrolyte and the surface of each of the electrodes, and (ii) modification of the cathode itself by doping one or more structurally-stabilizing element(s) into the lattice structure of the cathode, or by introducing one or more physical barrier layer(s) on the cathode surface to passivate unwanted side reactions.^{11,13,14} This passivation of the cathode has been known as the cathode coating of LIBs, which has been studied for more than a decade with numerous coating materials.^{9–12} Although traditional coating materials, such as Li-free electronic insulators (*e.g.*, Al_2O_3 , AlF_3 , MgO , ZrO_2 , SiO_2), have demonstrated their ability to enhance the overall performance of LIBs, their limited Li-ion conductivity and solubility result in increased overpotentials.¹⁶ This means that more energy is required for the coated electrode compared to that of its uncoated counterpart to achieve the same electrochemical performance. These limitations have driven researchers to explore alternative coating materials, including lithium-containing compounds such as LiAlO_2 ^{17,18} and LiBO_2 ,^{5–8} whose electronic and Li-ion conduction models resemble those of the SEIs of LIBs.¹⁶

LiBO_2 , in particular, has recently been found to have significant potential to stabilize cathode materials, especially some layered oxides LiMO_2 ($M = \text{Ni}_x\text{Mn}_{1-x-y}\text{Co}_y$ with $0 \leq x, y \leq 1$), resulting in their increased overall electrochemical performance. Zhang *et al.*⁸ synthesized LiBO_2 -coated $\text{LiNi}_{0.6}\text{Co}_{0.2}\text{Mn}_{0.2}\text{O}_2$ for use in polyetheracrylate-based solid-state batteries, reporting a capacity retention of 84.3% after 150 cycles at 0.5C. Wang *et al.*¹⁹ found that LiBO_2 coating on $\text{LiNi}_{0.5}\text{Co}_{0.2}\text{Mn}_{0.3}\text{O}_2$ significantly improved its performance, achieving a retention of 90.1% after 100 cycles at 1C (3.0–4.6 V), compared to 72.3% for the unmodified material. Du *et al.*²⁰ applied a wet chemical method to coat LiBO_2 on $\text{LiNi}_{0.8}\text{Co}_{0.1}\text{Mn}_{0.1}\text{O}_2$ (NCM811), with the modified sample delivering a capacity of 157.7 mAh g^{-1} and a retention of 82.1% after 100 cycles at 1C and 4.5 V, outperforming the pristine material, which only had a capacity of 96.0 mAh g^{-1} and a retention of 50.8%. Lim *et al.*²¹ employed a solution method to coat $\text{LiNi}_{0.8}\text{Co}_{0.15}\text{Al}_{0.05}\text{O}_2$ with LiBO_2 , resulting in an initial capacity of 173 mAh g^{-1} and a retention of 94.2% after 100 cycles at 1C and 55 °C (3–4.3 V), which was superior to the 75.3% retention of unmodified samples.

Interestingly, LiBO_2 material coating on the surface of the cathode and doping boron into its crystal lattice have been

reported to synergistically enhance the overall performance of LIBs even further.^{5,6} For instance, Gao *et al.*⁵ reported that doping the cathodes with boron and coating them with LiBO_2 simultaneously boosted the electrochemical performance of the $\text{LiNi}_{0.6}\text{Co}_{0.1}\text{Mn}_{0.3}\text{O}_2$ (NCM613) cathode material operating at high charging/discharging rates, high temperatures, and high voltage cutoffs. Specifically, they revealed that their optimized NCM613 cathode coated with LiBO_2 and doped with boron has shown its outstanding cycling stability at room temperature, retaining 94.8% of its capacity at 1C after 100 cycles and 70.7% at 5C after 1000 cycles (2.8–4.5 V). At 45 °C and 1C, this modified NCM613 cathode also exhibited an initial discharge capacity of 195.8 mAh g^{-1} and maintained 88.0% of its capacity after 100 cycles, significantly outperforming the pristine NCM613 material, which only retained 67.3%. Moreover, the modified NCM613 material exhibited a higher lithium diffusion coefficient compared to its pristine NCM613 counterpart. Therefore, this work suggested an alternative approach to modifying the cathode materials of LIBs using the B element. This suggestion is in line with published studies^{22,23} demonstrating that the recently-developed family of boron-containing liquid electrolytes resulted in a favorable SEI, stabilizing and enhancing the performance of LIB cathode materials.

Because of its aforementioned potential as an effective cathode coating, recent basic research focusing on ionic and electronic properties of LiBO_2 has been increasingly revived in order to provide insights into the design and development of more effective LiBO_2 -based cathode coatings. Hirose *et al.*²⁴ conducted an experimental study on Li-ion transport in high-density polycrystalline LiBO_2 polymorphs synthesized under high pressures, the tetragonal LiBO_2 (t-LBO) and the monoclinic LiBO_2 (m-LBO). It is noted that while the t-LBO is only stable at high temperatures and/or high pressures, the m-LBO is stable at ambient temperature and atmospheric pressure. Using electrochemical impedance spectroscopy (EIS) to characterize the ion transport in their samples between 450 and 520 K, they found that the intra-grain and total ionic conductivities of the t-LBO with a three-dimensional Li–Li network were between 10^{-6} and 10^{-5} S cm^{-1} , with total conductivity *ca.* 10^{-6} S cm^{-1} . These values were consistently higher than those observed for the m-LBO with a two-dimensional network across the entire temperature range. Islam *et al.*²⁵ conducted density functional theory (DFT) calculations for the formation and diffusion of Li defects in m-LBO crystals and the effect of Li defects on the electronic density of states (DOS). They found that the defect modified the DOS of the perfect crystal with the presence of defect levels in the band gap. They also found that there are numerous possible pathways of Li^+ diffusion in the m-LBO, including those along the *xy* plane.

In our previous work,²⁶ we systematically examined the formation of Li, B, and O vacancies in both m-LBO and t-LBO and their effects on Li-ion transport using DFT calculations. We found that the formation energy of lattice vacancies increases from Li to O, then to B. For Li-ion transport, O vacancies lower the migration energy barrier (E_m) in m-LBO but increase it in t-LBO. In contrast, B vacancies reduce E_m in



both m-LBO and t-LBO, enhancing diffusivity and ionic conductivity. Thus, our previous work suggested that creating B vacancies using an approach such as neutron irradiation²⁷ may improve ionic conductivity in the LiBO_2 material. Basalaev *et al.*²⁸ studied the electronic band structures of both m-LBO and t-LBO using DFT calculations and showed that the values of the band gap are 10.4 eV and 7.6 eV for t-LBO and m-LBO, respectively. Nevertheless, how the lattice vacancies affect the electronic band structures of these polymorphs is still an open question. Answering this question is of paramount importance for better designing and engineering conformal LiBO_2 -based surface coatings of the cathode.^{9–12} This is because a conformal surface coating is supposed to simultaneously hold multiple conflicting protective mechanisms of its functionalities, including ion and electron transport properties, which are still poorly understood.^{9–12,16}

In this work, we address the aforementioned question, aiming at the effects of lattice vacancies on the electronic band structures of both m-LBO and t-LBO (see Fig. 1) using DFT calculations, the setup of which is presented in Section 2. Section 3 presents and discusses the outcome of our DFT calculations. Although thematically related to our previous study,²⁶ the present work focuses on a distinct but equally critical aspect: the influence of vacancies on electron, rather than ion, transport. Our results reveal a strong phase dependence: boron vacancies in monoclinic LiBO_2 promote both ionic and electronic conduction, supporting its role as a cathode coating, while in tetragonal LiBO_2 , such vacancies increase electronic resistivity, reinforcing its suitability as a solid electrolyte. Whereas our findings provide practical insight into the phase-selective design of LiBO_2 for lithium-ion battery applications, additional criteria such as thermodynamic and electrochemical stability, delithiation resistance, *etc.*, are also important for practical applications and will require further study. Finally, since both

ionic and electronic conductivities govern the utility of materials such as solid-state electrolytes or electrode coatings, our current results complement the previous work to advance a more complete understanding of how lattice vacancies modulate charge transport in LiBO_2 polymorphs. Taken together, the two studies constitute a unified investigation into the dual functionality driven by defects of LiBO_2 in lithium-ion batteries.

2. Computational methods

DFT calculations were performed to investigate the electronic band structures of both pristine and defective crystals of the m-LBO (space group $P2_1/c$) and t-LBO (space group $I42d$) polymorphs of LiBO_2 materials. These calculations were carried out using the WIEN2K code,^{29,30} which is based on the full-potential linearized augmented plane wave (FP-LAPW) and local orbital basis set method to solve Kohn–Sham equations to calculate the electronic band structures of the LiBO_2 solid crystals.³¹

The exchange–correlation energy functional was treated using the Perdew–Burke–Ernzerhof (PBE) form of the generalized gradient approximation (GGA).³² All calculations were performed within an all-electron, self-consistent framework, with both core and valence electrons included. For both the tetragonal (t-LBO) and monoclinic (m-LBO) polymorphs of LiBO_2 , total energies were converged to within 0.0001 Ry.

The use of GGA-PBE in this study was motivated by two key considerations. First, employing the same GGA-PBE framework as in our previous study on ion transport in LiBO_2 ²⁶ ensures methodological consistency with that study, which needed more than 500 separate DFT calculations to evaluate the effects of Li, B, and O lattice vacancies on ion transport in the two t-LBO and m-LBO polymorphs. Performing these calculations with a hybrid functional such as HSE06, which typically incurs a computational cost 10 to 100 times higher due to the $O(N^4)$ scaling of exact exchange, would require resources far beyond those available to us. Second, although GGA-PBE underestimates the absolute band gap, it reliably captures qualitative trends in relative electronic structure modifications. As demonstrated in Section 3, our findings regarding the comparative electronic properties of m-LBO and t-LBO (*e.g.*, the narrower band gap of m-LBO) are consistent with previously reported results obtained from hybrid functional calculations.²⁸ Therefore, we regard the use of GGA-PBE as a judicious compromise between computational efficiency and predictive accuracy.

A Monkhorst–Pack k -point mesh of $5 \times 5 \times 5$ k -points was adopted for the Brillouin zone sampling and the Brillouin zone integration was carried out using the modified tetrahedron method,³³ and the plane wave cut-off parameter $R_{K\text{max}}$ was set to 7.00, ensuring adequate basis set completeness. The Muffin-tin radii were carefully chosen to avoid overlap and were respectively set to 1.61, 1.14, and 1.32 Bohr for Li, B, and O atoms in m-LBO and were respectively set to 1.66, 1.30, and 1.51 Bohr for Li, B, and O atoms in t-LBO.

In principle, one could computationally explore a broad set of defective crystal structures of the form $\text{Li}_{1-x}\text{B}_{1-y}\text{O}_{2-z}$, where

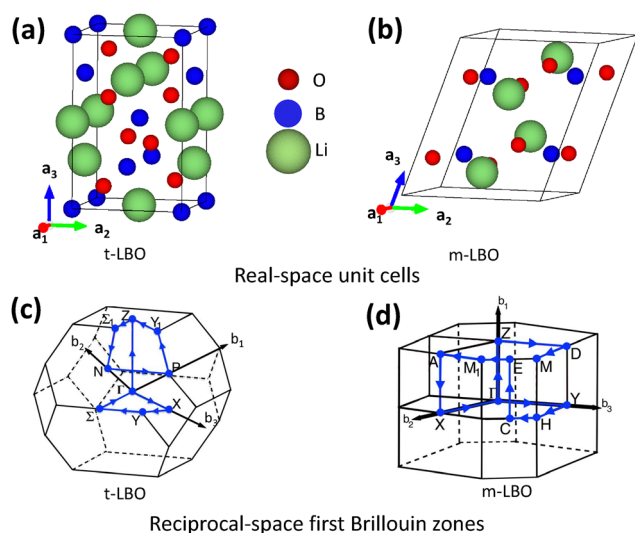


Fig. 1 Structures of LiBO_2 in real and reciprocal spaces. Unit cells of t-LBO (a) and m-LBO (b) polymorphs of LiBO_2 crystal. First Brillouin zone of t-LBO (c) and m-LBO (d).



$0 \leq x, y < 1$ and $0 \leq z < 2$, to capture a wide spectrum of vacancy concentrations and their potential effects on electronic transport properties. Such an extensive study would account for the diversity of defect configurations and enable a comprehensive understanding of their impact. However, in the present work, we limit our analysis to a selected set of representative cases for both polymorphs of LiBO_2 : (i) the pristine structure ($x = y = z = 0$); (ii) a Li vacancy ($x = 0.25, y = z = 0$); (iii) a B vacancy ($y = 0.25, x = z = 0$); and (iv) an O vacancy ($z = 0.125, x = y = 0$).

Rather than focusing on isolated point defects at dilute concentrations, our study investigates $1 \times 1 \times 1$ supercells in which vacancy defects are present at high concentrations, thereby allowing us to probe the regime of strongly interacting defects. This corresponds to a vacancy density on the order of $\sim 8 \times 10^{23} \text{ cm}^{-3}$, or one vacancy per unit cell, as previously described in ref. 26. While such concentrations may be significantly higher than those typically encountered in conventional materials processing, they are plausible under extreme service conditions, such as cathode coatings in lithium-ion batteries (LIBs) subjected to prolonged neutron irradiation.²⁶ These scenarios are relevant for LIB applications in harsh environments, including nuclear reactors and deep-space missions,^{34–37} where high radiation doses are expected to generate elevated concentrations of lattice vacancies.

Experimentally, neutrons, in particular thermal neutrons, can generate either Li or B vacancies in compounds containing these elements such as LiBO_2 by transmuting either the ^{10}B (19.9% of B) or ^6Li (7.5% of Li) isotopes at their respective lattice sites via the $^{10}\text{B} + {}^1_0n$ (0.025 eV) $\rightarrow$ ^4_2He (1.47 MeV) + ^7_3Li (0.84 MeV) + γ (0.48 MeV) or the $^6\text{Li} + {}^1_0n$ (0.025 eV) $\rightarrow$ ^3_1H (2.73 MeV) + ^4_2He (2.05 MeV) + 4.78 MeV nuclear reactions.^{38,39} As observed in the outcomes of our experimental characterization of the m-LBO surface irradiated with thermal neutrons at various neutron doses (not shown here, will be reported elsewhere⁴⁰), the concentration of lattice vacancies was found to be dose-dependent. Accordingly, for sufficiently high neutron doses, the vacancies were no longer isolated, but aggregated into clusters and mutually interacting. Thus, our $1 \times 1 \times 1$ supercells with one B vacancy per unit cell, which contains four B atoms, could model the case that all the ^{10}B sites are completely transmuted with intense and/or prolonged exposure of the material to thermal neutrons, *e.g.*, in the extreme environments mentioned above.

Defective lattice structures were modeled by introducing one charge-neutral lattice vacancy per supercell at either Li or B or O site into the respective supercells for each of the m-LBO and t-LBO polymorphs constructed from the fully optimized pristine (*i.e.*, perfect or nondefective) crystal structures, which were subject to full relaxation of atomic positions and lattice parameters until the force acting on each atom was less than 1 mRy Bohr⁻¹ (see ref. 26). Afterwards, the defective supercells were also fully relaxed until the aforementioned criteria of total-energy and force convergences were fulfilled. It is important to note that in battery applications the devices are connected to external circuitry, *i.e.*, to effective electron reservoirs which allow for dynamic charge neutrality. This justifies our assumption of

neutral defects. Additionally, by restricting our study to such neutral vacancies allows us to study their specific effect on the electronic structure without the need to consider long-range band structure effects of local charges. This approach allows us to focus solely on the effect of neutron-induced defect behavior, setting a solid foundation for future studies to consider the additional effect of charge states and defect combinations.

The electronic band structures of both pristine and defective crystals were calculated along high-symmetry k -paths of the first Brillouin zones as shown in Fig. 1c and d. The k -paths were created using the XCrysDen software.⁴² The chosen k -paths were selected to reflect the symmetry of the specific supercell structures (see ref. 41). In the presence of vacancies, our previous study²⁶ shows that while lattice vacancies modify the lattice constants a , b , and c , the values of α , β , and γ are still preserved (see Table 3 of ref. 26) to enable the defective lattice structure to have the same either monoclinic symmetry of the m-LBO polymorph (space group $P2_1/c$) or the tetragonal symmetry of the t-LBO polymorph (space group $I4_2d$) as their pristine counterparts. For each defective structure, we verified the preservation of its symmetry using the XCrysDen software.

Specifically, for the t-LBO polymorph of a centered tetragonal lattice structure ($\mathbf{a}_1 = (a, 0, 0)$, $\mathbf{a}_2 = (0, b, 0)$, and $\mathbf{a}_3 = (0, 0, c)$, $a = b < c$, $\alpha = \beta = \gamma$), the k -path (see Fig. 1c) is $\Sigma \rightarrow \Gamma \rightarrow X \rightarrow Y \rightarrow \Sigma \rightarrow \Gamma \rightarrow Z \rightarrow \Sigma_1 \rightarrow N \rightarrow P \rightarrow Y_1 \rightarrow Z$, of which the fraction coordinates of the k -points are $\Sigma(-\eta, \eta, \eta)$, $\Gamma(0, 0, 0)$, $X(0, 0, \frac{1}{2})$, $Y(-\zeta, \zeta, \frac{1}{2})$, $Z(\frac{1}{2}, \frac{1}{2}, -\frac{1}{2})$, $\Sigma_1(\eta, 1 - \eta, -\eta)$, $N(0, \frac{1}{2}, 0)$, $P(\frac{1}{4}, \frac{1}{4}, \frac{1}{4})$, and $Y_1(\frac{1}{2}, \frac{1}{2}, \eta)$ given the reciprocal lattice vectors being respectively $\mathbf{b}_1(\frac{2\pi}{a}, 0, \frac{2\pi}{a})$, $\mathbf{b}_2(\frac{2\pi}{a}, \frac{2\pi}{a}, 0)$, and $\mathbf{b}_3(0, \frac{2\pi}{c}, \frac{2\pi}{c})$, where $\eta = \frac{1+a^2/c^2}{4}$ and $\zeta = \frac{a^2}{2c^2}$ are determined from the corresponding lattice parameters a (note that $a = b$ for a tetragonal lattice) and c of the corresponding pristine or defective tetragonal crystals (see ref. 26).

In contrast, for the m-LBO polymorph of a monoclinic lattice structure ($\mathbf{a}_1 = (a, 0, 0)$, $\mathbf{a}_2 = (0, b, 0)$, and $\mathbf{a}_3 = (0, c \cos \alpha, c \sin \alpha)$, where $a, b < c$, $\alpha < 90^\circ$, $\beta = \gamma = 90^\circ$), the k -path (see Fig. 1d) is $\Gamma \rightarrow Y \rightarrow H \rightarrow C \rightarrow E \rightarrow M_1 \rightarrow A \rightarrow X \rightarrow \Gamma \rightarrow Z \rightarrow D \rightarrow M$ with the fractional coordinates of the high-symmetry k -points being $\Gamma(0, 0, 0)$, $Y(0, 0, \frac{1}{2})$, $H(0, \eta, 1 - \nu)$, $C(0, \frac{1}{2}, \frac{1}{2})$, $E(\frac{1}{2}, \frac{1}{2}, \frac{1}{2})$, $M_1(\frac{1}{2}, 1 - \eta, \nu)$, $A(\frac{1}{2}, \frac{1}{2}, 0)$, $X(0, \frac{1}{2}, 0)$, $Z(\frac{1}{2}, 0, 0)$, $D(\frac{1}{2}, 0, \frac{1}{2})$, and $M(\frac{1}{2}, \eta, 1 - \nu)$ given the reciprocal lattice vectors being respectively $\mathbf{b}_1(\frac{2\pi}{a}, 0, 0)$, $\mathbf{b}_2(0, \frac{2\pi}{b}, \frac{2\pi \cot \alpha}{b})$, and $\mathbf{b}_3(0, 0, \frac{2\pi}{c \sin \alpha})$, where $\eta = \frac{2-(b/c)\cos(\alpha)}{2\sin^2(\alpha)}$ and $\nu = \frac{1}{2} - \eta \frac{c}{b} \cos \alpha$ are determined from the corresponding lattice parameters a , b , c , and α of the corresponding pristine or defective monoclinic crystals (see ref. 26).

3. Results and discussions

This section focuses on presenting and discussing the outcome of our DFT calculations of electronic band structures for both t-LBO (Section 3.3) and m-LBO (Section 3.2) polymorphs. Importantly, the present study does not rely solely on the



analysis of the density of states to infer charge transport mechanisms. Instead, our approach is based on a detailed examination of the modifications in the electronic band structure—specifically, the energy dispersion relationship $E(k)$ near the conduction band minimum (CBM), valence band maximum (VBM), and any defect-induced mid-gap states. From these dispersion characteristics, we extract qualitative insights into the effective mass (*via* band curvature) and the effective band gap, both of which are directly related to carrier mobility and intrinsic carrier concentration, respectively. These quantities, in turn, govern the temperature-dependent electronic conductivity. Recall that the electron mobility μ is given by $\mu = \frac{e\tau}{m^*}$, where e is the elementary charge, τ is the carrier scattering time, and m^* is the effective mass. The effective mass is derived from the curvature of the band structure *via* $\frac{1}{m^*} = \frac{1}{\hbar^2} \frac{d^2 E(k)}{dk^2}$, where $E(k)$ is the energy dispersion relationship and $\hbar$ is the reduced Planck constant. The conductivity is $\sigma = en\mu$, where $n \propto \exp(-E_g/k_B T)$ is the carrier concentration, which is a function of temperature and is governed by the band gap (see ref. 43). Although we do not explicitly calculate numerical values for conductivity or mobility, our analysis provides qualitative insight into how lattice vacancies influence electronic transport. Moreover, by extracting band gap values from the calculated band structures, we estimate the relative changes in carrier concentrations (electrons and holes) at room temperature, the most relevant operational temperature for lithium-ion batteries. These estimates, while approximate, allow us to assess trends in electronic behavior across different defect configurations and polymorphs of LiBO_2 .

3.1 Effects of lattice vacancies on the electronic band structure of tetragonal LiBO_2

Fig. 2 shows the electronic band structures calculated for the tetragonal LiBO_2 polymorph: pristine (Pri t-LBO, Fig. 2a), Li-vacancy (Livac t-LBO, Fig. 2b), B-vacancy (Bvac t-LBO, Fig. 2c), and O-vacancy (Ovac t-LBO, Fig. 2d) supercells.

Pristine tetragonal LiBO_2 (Pri t-LBO). Fig. 2a presents the calculated electronic band structure of the Pri t-LBO, which manifests an indirect band-gap electronic band structure with the valence band maximum (VBM) located at the N k -point of the Fermi level and the conduction band minimum (CBM) located at the Γ k -point and separated from the VBM by a band gap of $E_g = 7.3$ eV, which is less than that ($E_g = 10.4$ eV) calculated by Basalae *et al.*²⁸ using a computationally demanding hybrid functional method. This difference in the values of the band gap E_g is attributed to the limitation of a less computationally demanding GGA method used in our current work to underestimate the value of E_g .

It is noted that, (1) the energy level of the top of the valence band at the Γ and the Z k -points are very close (≈ 0.2 eV) to the Fermi level, and (2) the energy levels of the bottom of the conduction band at Γ and Z k -points are also very close (≈ 0.2 eV). Therefore, the cross-band-gap Γ -to- Γ or Z-to-Z excitations of the electron from the top of the valence band

to the bottom of the conduction band might need energies approximately of the order of $E_g = 7.3$ eV without the need of a phonon absorption to fulfill the law of momentum conservation, suggesting that the Pri t-LBO might behave in reality as a quasi-indirect band gap insulator at elevated temperatures under a stimulation of sufficiently strong electric potentials or an absorption of extreme/deep ultraviolet (EUV/DUV) photons.

Finally, regarding the electric properties of charge carriers, Fig. 2a shows that the curvatures of the lowest energy dispersion relationship $E(k)$ curve of the conduction band at Γ and Z k -points are quite similar and are both larger than those of the highest $E(k)$ curve of the valence band at these k -points, suggesting that the effective mass of an electron at the bottom of the conduction band is less than that of an electronic hole at the top of the valence band, and hence the electron, if excited to the conduction band, would likely be more mobile than the electronic hole that is excited to the valence band.

Li-vacancy tetragonal LiBO_2 (Livac t-LBO). Fig. 2b shows the calculated band structure of the Livac t-LBO. At first glance, it appears that there are no obvious modifications of the band structure of the t-LBO polymorph in the presence of Li vacancies. However, a closer look shows the following important changes.

First, while the minima of the lowest energy dispersion relationship $E(k)$ curve of the conduction band are still located at Γ and Z k -points with the VBM at the former, all of the $E(k)$ curves of the conduction band were shifted up in energy by ≈ 0.5 eV, a considerable amount of energy compared to the thermal energy (≈ 0.026 eV).

Second, the curvatures of the lowest $E(k)$ curve of the conduction band at the Γ and Z k -points of the Livac t-LBO (see Fig. 2b) are both smaller than those of the Pri t-LBO (see Fig. 2a), indicating that the Li vacancies reduce the mobility of electrons excited to the conduction band of the m-LBO polymorph.

Third, the CBM of the band structure of the Livac t-LBO is still located at the same k -point (*i.e.*, the N k -point) as that of the Pri t-LBO, suggesting that the indirect band gap feature is still preserved with the presence of Li vacancies. The Fermi level, however, is shifted down by ≈ 0.2 eV to the energy levels of the maxima of the highest $E(k)$ of the conduction band at the Γ and Z k -points. This means that the energy levels introduced by the Li vacancies between VBM and the Fermi level manifest the acceptor levels of the Li vacancies, and the Livac t-LBO might behave as a p-type semiconductor with a wide band gap of $E_g = 7.8$ eV, which is 0.5 eV higher than that ($E_g = 7.3$ eV) of the Pri t-LBO as a result of the uplifting of the conduction band by 0.5 eV as aforementioned.

Fourth, the curvatures of the highest $E(k)$ of the valence band at its maxima (*i.e.*, at the Γ , Z, and N k -points) are smaller (*i.e.*, the $E(k)$ curve at these k -points is flatter for the Livac t-LBO than for the Pri t-LBO), suggesting that electronic holes (which are excited to the valence band) possess increased effective masses, and hence reduced mobilities, in the presence of Li vacancies in the t-LBO polymorph.



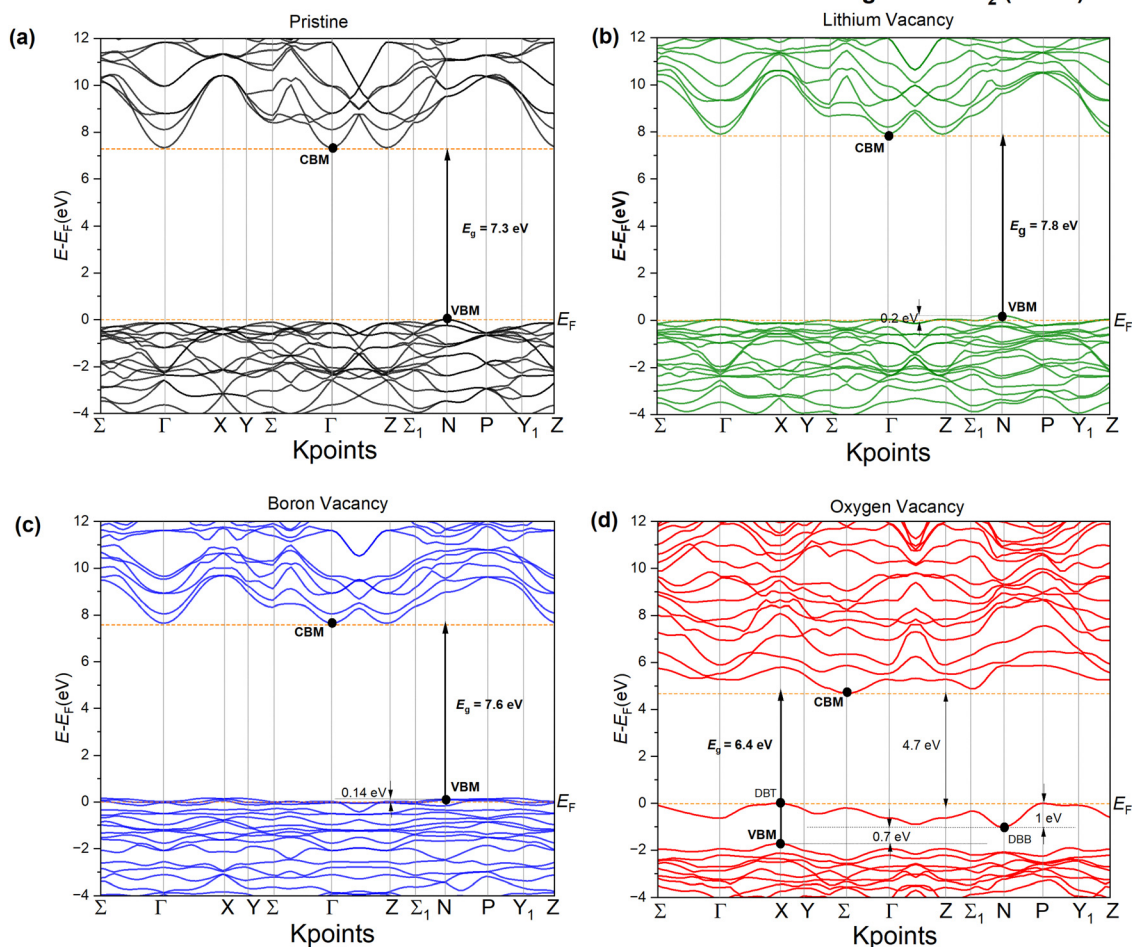
Effects of Vacancies on Electronic Band Structures of Tetragonal LiBO_2 (t-LBO)

Fig. 2 Calculated electronic band structure of tetragonal LiBO_2 (t-LBO) for (a) pristine, (b) Li-vacancy, (c) B-vacancy, and (d) O-vacancy supercells. The conduction band minimum, valence band maximum, defect band top, and defect band bottom are abbreviated as CBM, VBM, DBT, and DBB, respectively.

Finally, let us now consider the application perspective of the Livac t-LBO. In our previous work (see ref. 26), we have shown that Li vacancies need to be present in t-LBO crystals for Li-ion transport *via* a vacancy-mediated diffusion mechanism. Of all the 9 Li-ion diffusion pathways investigated for the Livac t-LBO, the lowest diffusion barrier is $E_m = 0.46$ eV and thus the estimated ionic conductivity is $\sigma_{\text{ion}} = 7.0 \times 10^{-5} \text{ S cm}^{-1}$. Nevertheless, this value of σ_{ion} and the band gap $E_g = 7.8$ eV both show that the Livac t-LBO is not suitable as a potential conformal cathode coating of Li-ion batteries since the coating should be both a good ionic conductor and an electronic conductor.

B-vacancy tetragonal LiBO_2 (Bvac t-LBO). Fig. 2c shows the band structure of the Bvac t-LBO. In general, the modification of the band structure induced by B vacancies exhibits similar features to those discussed above for the Livac t-LBO with some minor differences. These features are briefly described as follows.

First, the conduction band of the Bvac t-LBO is shifted up by 0.3 eV, which is less than that (0.5 eV) of the Livac t-LBO while the minima of the lowest $E(k)$ curve of the conduction band are

still located at both the Γ and Z k -points with the CBM located at the Γ k -point.

Second, the curvatures of the lowest $E(k)$ curve of the conduction band at both the Γ and Z k -points are smaller than those of the Livac t-LBO and the Pri t-LBO, meaning that the mobility of the electron excited to the conduction band is reduced even further in the presence of B vacancies in the t-LBO polymorph.

Third, like the case of the Livac t-LBO, the VBM is located at the N k -point while the Fermi level is pushed down by 0.14 eV, which is slightly less than that of the Livac t-LBO, meaning that the energy band of the acceptor introduced by B vacancies, which are located between the Fermi level and the VBM (see Fig. 2c), is a bit narrower than that introduced by Li vacancies (see Fig. 2b). Yet, the Bvac t-LBO might still behave as a p-type semiconductor with a wide band gap of $E_g = 7.6$ eV, which is wider than that ($E_g = 7.3$ eV) of the Pri t-LBO, but narrower than that ($E_g = 7.8$ eV) of the Livac t-LBO.

Fourth, the highest $E(k)$ curve of the valence band of the Bvac t-LBO is flatter than that of the Livac t-LBO, meaning that



the electronic hole of the Bvac t-LBO is heavier with its smaller mobility.

Finally, let us now compare the applicability of the Bvac t-LBO and the Livac t-LBO as a solid electrolyte or a conformal cathode coating of Li-ion batteries. In our previous DFT work (see ref. 26), we reported that the Bvac t-LBO has an ionic conductivity ranging from $\sigma_{\text{ion}} = 2.5 \times 10^{-3} \text{ S cm}^{-1}$ to $\sigma_{\text{ion}} = 7.5 \text{ S cm}^{-1}$, which is in the upper range of values of the ionic conductivity of the currently developed solid electrolytes for all-solid Li-ion batteries and shows that the Bvac t-LBO is promising, in terms of its Li-ion transport, for its applications as both a solid electrolyte and a conformal cathode coating (*i.e.*, a good ionic conductor). However, in terms of the electron/hole transport deduced from our DFT-calculated electronic band structure shown in Fig. 2c, the combined findings of our previous²⁶ and current studies suggest that the Bvac t-LBO would be a potential candidate as a solid electrolyte rather than a conformal cathode coating.

This is because the Bvac t-LBO possesses a wide band gap ($E_g = 7.6 \text{ eV}$) (*i.e.*, a good electronic insulator) which exceeds the high-voltage window (3 to 6 V) of Li-ion batteries. In addition, the acceptor levels are located very close to the top of the valence band that will accept electrons to the sites of B vacancies, which are then negatively charged to facilitate Li-ion transport *via* a vacancy-mediated diffusion mechanism (*i.e.*, positively-charged Li^+ ions could be electrostatically driven by the negatively-charged B vacancies to hop into their sites on their diffusion pathways²⁶). Yet, a good electronic insulator is not suitable as a conformal cathode coating since there still is a certain voltage drop across the coating layer due to a high electronic resistivity.

O-vacancy tetragonal LiBO_2 (Ovac t-LBO). Fig. 2d presents the band structure of the Ovac t-LBO. Compared to Li and B vacancies in the t-LBO polymorph, O vacancies modify the band structure significantly.

First, the conduction band of the Ovac t-LBO is substantially shifted down in energy by 2.6 eV, instead of shifting up as seen for both the Livac t-LBO (up by 0.5 eV) and the Bvac t-LBO (up by 0.3 eV).

Second, the shapes of all energy dispersion $E(k)$ curves of the conduction band are completely changed in comparison with those of the Pri t-LBO. The minima of the lowest $E(k)$ curve of the conduction band are moved from the Γ and Z k -points for the Pri t-LBO to the Σ and Σ_1 k -points with the CBM located at the Σ k -point for the Ovac t-LBO. The curvature of the lowest $E(k)$ curve around the Σ k -point (*i.e.*, at the CBM) for the Ovac t-LBO is the least among the four systems shown in Fig. 2, meaning that the effective masses of electrons in the Ovac t-LBO are the largest and their mobilities are lowest among electrons in each of the four systems of the t-LBO polymorph, whose electronic band structures are shown in Fig. 2.

Third, the valence band is also substantially modified. The VBM is now moved from the N k -point for the Pri t-LBO to the X k -point for the Ovac t-LBO. In addition, the VBM is located at the energy level of 1.7 eV below the Fermi level. The highest $E(k)$ curve of the valence band is modified in such a way that it is more wavy than those of the Livac t-LBO and Bvac t-LBO. The

energy band gap ($E_g = 6.4 \text{ eV}$) is thus less than that for the other three systems.

Fourth, a remarkable new feature is generated: deep defect energy levels are introduced by oxygen vacancies into the band gap. The bottom of the defect band (DBB) and the top of the defect band (DBT) are respectively located at the N k -point (0.7 eV above the VBM) and at the X and P k -points (4.7 eV below the CBM). The vacancy sites of O^{2-} anions act as the deep centers to trap electrons. Thus, they were called in the fields of electro- and photo-optics as the color, Forbe, or F-centers.⁴⁵ An F center is a type of crystallographic defect that forms when a negative ion is removed from a crystal lattice, leaving behind a vacancy. This vacancy can capture an electron, potentially giving the material a colored appearance.

Finally, because of electrons being trapped to the deep defect levels of oxygen vacancies, the Fermi energy level is no longer located at the VBM, but rather moves to the top of the defect band (TDB), resulting in a reduced effective indirect band gap, which is as small as $E_g^{\text{eff}} = 4.7 \text{ eV}$, indicating that the Ovac t-LBO is not as good an electronic insulator as the Bvac t-LBO for use as a solid electrolyte. In addition, our previous study (see ref. 26) showed that the ionic conductivity of the Ovac t-LBO is estimated to be $\sigma_{\text{ion}} = 7.6 \times 10^{-8} \text{ S cm}^{-1}$, which is too low to be a good ionic conductor and inhibits the Ovac t-LBO from being either a solid electrolyte or a conformal cathode coating.

3.2 Effects of lattice vacancies on the electronic band structure of monoclinic LiBO_2

Fig. 3 shows the calculated electronic band structures for the monoclinic LiBO_2 polymorph: pristine (Pri m-LBO, Fig. 3a), Li-vacancy (Livac m-LBO, Fig. 3b), B-vacancy (Bvac m-LBO, Fig. 3c), and O-vacancy (Ovac m-LBO, Fig. 3d) supercells.

Pristine monoclinic LiBO_2 (Prim-LBO). As one can see in the Fig. 3a, different from the Pri t-LBO discussed above in Section 3.3, the electronic band structure of the Pri m-LBO exhibits a direct band gap of $E_g = 5.8 \text{ eV}$. As a typical descriptor of a direct-band-gap intrinsic insulator, both the conduction band minimum (CBM) and valence band maximum (VBM) are located at the same Γ k -point, and the Fermi level is aligned with the VBM. In addition, the value of E_g is also less than the value ($E_g = 7.6 \text{ eV}$) calculated by Basalae *et al.*²⁸ using an expensive hybrid functional method. Yet, similar to their calculations, our calculations showed that the value of E_g of the Pri m-LBO is smaller than that of the Pri t-LBO, even though both were underestimated due to the limitation of the GGA approximation we used in this work. Finally, it is suggested that while both the Pri m-LBO and the Pri t-LBO are good electronic insulators, the latter apparently insulates electrons even better.

Li-vacancy monoclinic LiBO_2 (Livac m-LBO). The electronic band structure of the Livac m-LBO is slightly modified from that of the Pri m-LBO in the presence of Li lattice vacancies, as shown in Fig. 3b.

First, the material retains a direct band gap at the Γ k -point, but the band gap energy is slightly reduced to $E_g = 5.6 \text{ eV}$, leading to an increased electron concentration, $n_e(T)$, at temperature T since $n_e(T) \propto e^{-E_g/k_B T}$.⁴³ Specifically, it is shown that



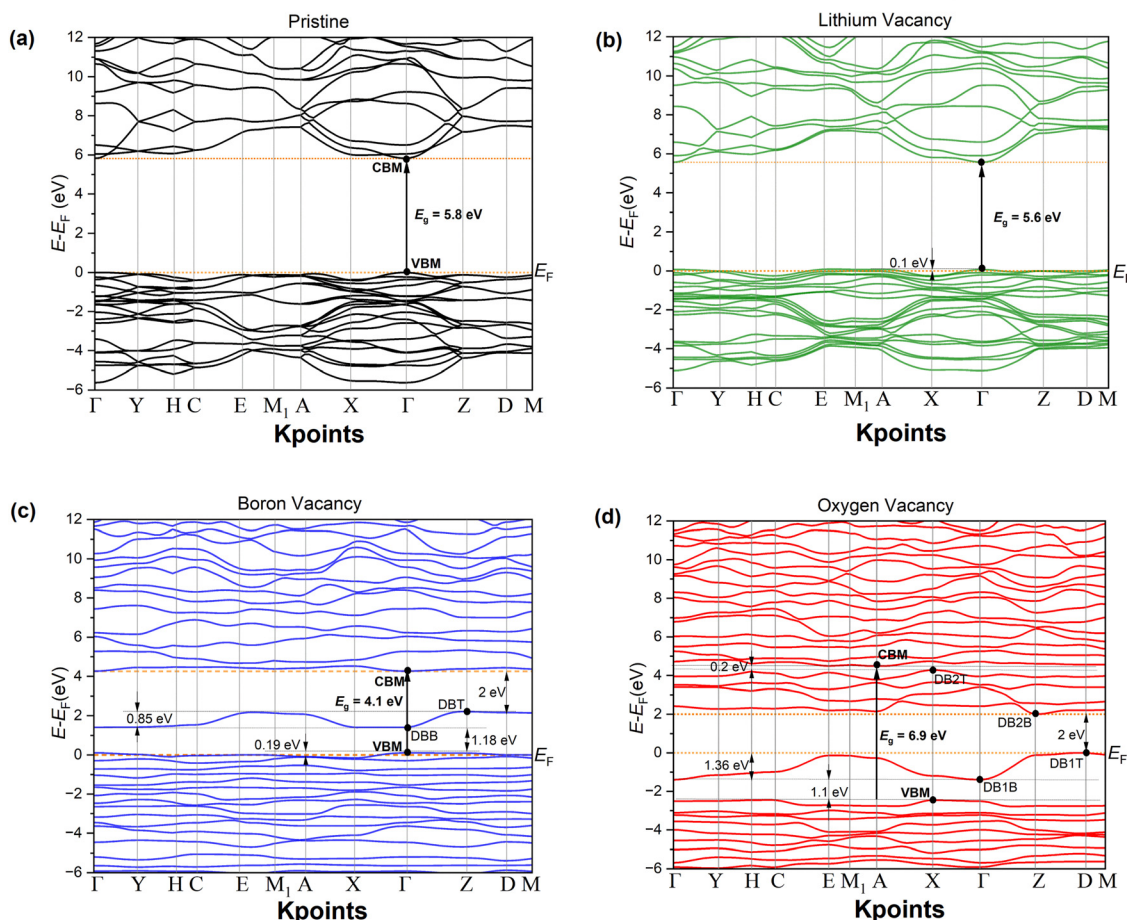
Effects of Vacancies on Electronic Band Structures of Monoclinic LiBO₂ (m-LBO)

Fig. 3 Calculated electronic band structure of monoclinic LiBO₂ (m-LBO) for (a) pristine, (b) Li-vacancy, (c) B-vacancy, and (d) O-vacancy supercells. The conduction band minimum, valence band maximum, defect band 1 top, defect band 1 bottom, defect band 2 top, and defect band 2 bottom are abbreviated as CBM, VBM, DB1T, DB1B, DB2T, and DB2B, respectively.

at the room temperature ($T = 300 \text{ K} = 0.026 \text{ eV}$) $n_e^{\text{Livac m-LBO}} / n_e^{\text{Pri m-LBO}} \approx e^{-5.6 \text{ eV} / 0.026 \text{ eV}} / e^{-5.8 \text{ eV} / 0.026 \text{ eV}} \approx 2.2 \times 10^3$. This estimated ratio suggests that n_e of the m-LBO polymorph might be increased by 1000 times at 300 K in the presence of Li vacancies.

Second, it is noted that the lowest energy dispersion relationship curve $E(k)$ of the conduction band near the CBM becomes flatter, indicating a slightly reduced curvature of $E(k)$ vs. k at the Γ k -point and consequently an increased electron effective mass and hence a reduced electron mobility.

Third, the Fermi level shifts down by 0.2 eV below the VBM, and the CBM lies at a slightly lower energy compared to the pristine structure. The defect energy states introduced by Li vacancies are shallowly positioned between the VBM and Fermi level, indicating that the Li vacancies act as electron acceptors and therefore the Livac m-LBO resembles a wide band-gap p-type semiconductor.

Finally, to assess the Livac m-LBO's applicability, we recall its Li-ion transport properties from ref. 26: its ionic conductivity is $\sigma_{\text{ion}} = 7.0 \times 10^{-13} \text{ S cm}^{-1}$, seven orders of magnitude lower than the $\sigma_{\text{ion}} = 7.0 \times 10^{-5} \text{ S cm}^{-1}$ of the Livac t-LBO. While both

are lower than the ionic conductivity of typical Li-ion battery liquid electrolytes ($\sigma_{\text{ion}} \sim 10^{-2}$ to 10 S cm^{-1} (ref. 44)), our comparison suggests that the Livac t-LBO is a better choice for a solid electrolyte due to its superior ionic conduction and electron insulation.

B-vacancy monoclinic LiBO₂ (Bvac m-LBO). Unlike Li vacancies discussed above, B vacancies in the m-LBO lattice modify its electronic band structure substantially.

First, Fig. 3c shows that the bottom of the conduction band of the Bvac m-LBO shifts down in energy by $\sim 2 \text{ eV}$ relative to that of the Pri m-LBO while the VBM is still located at the center of the first Brillouin zone (*i.e.*, the Γ k -point). Nevertheless, the curvature of the lowest energy dispersion relationship $E(k)$ curve of the conduction band at the CBM (*i.e.*, at the Γ k -point) is substantially reduced in comparison to those of both the Pri m-LBO and the Livac m-LBO, revealing a larger effective mass, and hence a smaller electron mobility, than those of both the Pri m-LBO and the Livac m-LBO.

Second, the VBM is still at the Γ k -point, but 0.2 eV above the Fermi level, indicating a direct band gap of $E_g \approx 4 \text{ eV}$, significantly smaller than those of both the Pri m-LBO and



Livac m-BO. Like the Livac m-LBO, the empty energy levels between the Fermi level and the VBM, which are introduced into the band gap by B vacancies, manifest shallow acceptor levels, typically occurring in the electronic band structure of a p-type wide band-gap semiconductor.

Third, there is a $E(k)$ curve in the middle region of the band gap. This new curve is attributed to the presence of B-vacancy-induced deep levels that collectively form a defect band of 0.85 eV in width, whose bottom-most level (abbreviated as DBB) is at the Γ k -point and is located at the energy level 1.18 eV above the VBM level and whose top-most level (abbreviated as DBT) is at the Z k -point and is located 2 eV below the CBM level. Unlike shallow levels, deep levels are located significantly far from the conduction or valence band edges and have highly localized wave functions, meaning the trapped charge carriers (in this case, the carriers are electronic holes rather than electrons as in the Ovac m-LBO discussed below) are tightly bound to defect sites (*i.e.*, B vacancies). At zero temperature, these deep levels are unoccupied by electrons (*i.e.*, empty states), suggesting that B vacancies may act as deep centers that trap electronic holes instead of electrons (*i.e.*, hole trapping centers) as shown in Fig. 3c. However, at a finite temperature, such as the room temperature at which electric devices like Li-ion batteries normally operate, these deep centers can act differently, depending on the amount of energy, ΔE , electrically (*i.e.*, via a bias dc electric field) or optically (*i.e.*, via excitation photons) imparted to electrons. If $\Delta E < 1.18$ eV B vacancies behave as hole-trapping centers. In contrast, if $1.8 \text{ eV} \leq \Delta E < 2$ eV they act as electron-trapping centers, negatively-charged B vacancies thus can attract positively-charged Li^+ ions that are on their pathways to transport toward the vacancies, facilitating Li-ion diffusion via a vacancy-mediated mechanism as reported in our previous work.²⁶ However, if $\Delta E \geq 2$ eV, they become recombination centers for both electrons and electronic holes, playing the role of stepping stones for charge carriers (*e.g.*, Li^+ ions, electrons, and electronic holes) to jump between sites in the Bvac m-LiBO lattice, facilitating the carriers to conduct an electric current upon the application of a sufficiently strong bias voltage (*e.g.*, the voltage across a LiBO_2 solid electrolyte sandwiched between the cathode and anode of Li-ion battery cells operating at voltages higher than 3 V), or facilitating a photocurrent upon the absorption of UV-vis photons. For $\Delta E \geq 2$ eV, both electrons and electronic holes can participate in the electric conduction; however, electronic holes are the major carriers and electrons are the minor carriers. This is because the number of electronic holes per unit volume (*i.e.*, the concentration), $n_h(T) \propto e^{-(E_{\text{DBB}}-E_{\text{VBM}})/k_B T} = e^{(-1.18\text{eV}/k_B T)}$, is much greater than that of electrons, $n_e(T) \propto e^{-(E_{\text{CBM}}-E_{\text{DBT}})/k_B T} = e^{-2\text{eV}/k_B T}$. For example, at room temperature ($T = 300 \text{ K} = 0.026 \text{ eV}$), $n_h/n_e \approx 4.6 \times 10^{13}$, meaning that the concentration of electronic holes is 13 orders of magnitude greater than that of electrons in the presence of B vacancies in the lattice of the m-LBO polymorph.

Finally, from the application points of view, in order to select the Bvac m-LBO as a solid electrolyte (a good electronic insulator) or/and a conformal cathode coating (a good

electronic conductor), one must first consider the electron transport with an effective band gap for the electron excitation at a finite temperature. This effective band gap would be estimated as $E_g^{\text{eff}} = E_{\text{CBM}} - E_{\text{DBT}} = 2 \text{ eV}$ (see Fig. 3c), which is much smaller than those of the Pri m-LBO ($E_g = 5.8 \text{ eV}$) and the Livac m-LBO ($E_g = 5.8 \text{ eV}$). Effectively, the excitation of electrons from the “stepping stone” energy level of the defect band top (DBT) at the Z k -point to the CBM at the Γ k -point manifests an indirect-band scenario. In addition, as reported in our previous work,²⁶ among the defective systems of the m-LBO polymorph considered in this section (Section 3.2), the Bvac m-LBO possesses the highest Li-ionic conductivity ($\sigma_{\text{ion}} = 2.8 \times 10^{-2} \text{ S cm}^{-1}$), which is about the lower bound of the ionic conductivity values of typical commercialized liquid electrolytes of Li-ion batteries ($\sigma_{\text{ion}} \sim 10^{-2}$ to 10 S cm^{-1} (ref. 44)). Overall, it is suggested that the Bvac m-LBO might be a promising candidate for the conformal cathode coating of Li-ion batteries (a good Li-ionic conductor and a good electronic conductor as well), not as a solid electrolyte (a good Li-ionic conductor and a good electronic insulator).

O-vacancy monoclinic LiBO_2 (Ovac m-LBO). Fig. 3d shows that O vacancies substantially modify the electronic band structure of the m-LBO polymorph.

First, similar to the Ovac t-LBO (see Fig. 2d), the VBM of the band structure of the Ovac m-LBO is moved from the Γ k -point to the X k -point. In contrast, this movement does not happen for both the Livac m-LBO (see Fig. 3b) and Bvac m-LBO (see Fig. 3c) for which the VBM was still located at the same Γ k -point as observed for the Pri m-LBO (see Fig. 3a).

Second, like the Ovac t-LBO, the energy level of the VBM of the Ovac m-LBO is not at the Fermi level, but about 2.46 eV below it. This is in contrast to the Bvac m-LBO (see Fig. 3c) for which the Fermi level is aligned with the VBM.

Third, similar to the Ovac t-LBO (see Fig. 2d) and the Bvac m-LBO (see Fig. 3c), there is also a curve of the energy dispersion relationship $E(k)$ in the middle of the band gap of the Ovac m-LBO. This curve is attributed to one of the defect bands (a deep-level defect band or defect band 1, DB1), which is introduced by O vacancies and spans an energy range of 1.36 eV from the defect band bottom (DB1B) (which is located at the Γ k -point and 1.1 eV above the VBM) to the defect band top (DB1T) (which is 2 eV below the bottom of another defect band introduced by oxygen vacancies). Similar to the Ovac t-LBO (see Fig. 2d), the Fermi level is located at the DB1T rather than at the VBM, meaning that these deep levels also act as the electron trapping centers or F centers. However, unlike the band structure of the Ovac t-LBO, there is another defect band, the defect band 2 (DB2), introduced into the band gap of the band structure of the Ovac m-LBO, just 0.2 eV below the CBM. The DB2 acts as an effective conduction band while the DB1 acts as an effective valence band. Thus, the band gap is effectively reduced to $E_g^{\text{eff}} = 2 \text{ eV}$ from the nominal band gap of $E_g = 6.9 \text{ eV}$. As a result, similar to the Bvac m-LBO, the Ovac m-LBO could be considered as a good electronic conductor for its use as a conformal cathode coating with the condition that its ionic conductivity must be high enough. However, our previous work



reported that the Ovac m-LBO is a poor Li-ion conductor with the critically low ionic conductivity of $\sigma_{\text{ion}} = 6.1 \times 10^{-6} \text{ S cm}^{-1}$ (ref. 26) compared to the ionic conductivity values of typical commercialized liquid electrolytes of Li-ion batteries ($\sigma_{\text{ion}} \sim 10^{-2}$ to 10 S cm^{-1} (ref. 44)). Overall, the Ovac m-LBO should be neither a solid electrolyte nor a conformal cathode coating of Li-ion batteries.

3.3 Summary

In summary, the present study demonstrates that the effects of lattice vacancies on electronic conduction in LiBO_2 are strongly phase dependent. Specifically, the introduction of boron vacancies in the monoclinic polymorph (m-LBO) enhances both ionic and electronic transport, whereas the same defect in the tetragonal polymorph (t-LBO) leads to increased electronic resistivity. Together with our previous investigation of the effects of lattice vacancies on ion transport in these materials,²⁶ these findings suggest that B-vacancy-defected t-LBO (Bvac t-LBO) may serve as a promising candidate for use as a solid electrolyte in lithium-ion batteries. In contrast, Bvac m-LBO appears well suited as a conformal cathode coating, where both ionic and electronic conduction could be beneficial under certain battery operating conditions. It must be stressed that, for the practical selection of functional materials, other additional factors and criteria, such as the interfacial, thermodynamic, and electrochemical stabilities, must also be evaluated.

Importantly, these phase-dependent behaviors are non-intuitive and represent a new insight into the defect-engineered functionality of LiBO_2 . Although our interpretations are grounded in well-established principles of semiconductor physics, the conclusions drawn here are specific, nontrivial, and directly relevant to materials design strategies for advanced lithium-ion battery technologies. Specifically, defect-induced deep- or mid-gap states (e.g., from B and O vacancies) can impact carrier transport by trapping or recombining carriers, depending on their energy levels and external conditions like electrochemical potential. This study examines how such states alter band edges and Fermi levels, particularly at operational voltages (e.g. 3–6 V in high-voltage Li-ion batteries). These states may become electronically active, influencing ionic conduction *via* carrier-assisted mechanisms. In an open-boundary electrochemical setup, shallow- to mid-gap defects can act as donors or acceptors, while deeper ones typically act as recombination centers. More detailed interface modeling is reserved for future work.

4. Conclusions

In conclusion, our studies highlight the potential of B vacancies in the LiBO_2 material to enhance its performance in lithium-ion battery applications. By improving the electronic insulation of the t-LBO polymorph and the electronic conduction of the m-LBO polymorph while boosting the ionic conductivity of both polymorphs, B lattice vacancy engineering emerges as a promising strategy for optimizing the properties of LiBO_2 as either a solid electrolyte (t-LBO) or a conformal cathode coating (m-LBO) for its

use in Li-ion batteries. In reality, as aforementioned in the introduction, m-LBO rather than t-LBO is stable at ambient temperature and pressure, and it is therefore suggested from our studies that the LiBO_2 material in the presence of B vacancies would be considered a potential conformal cathode coating in the form of crystalline films. It is noted that cathode coatings were reported to be partially or fully amorphous, though their structures largely depend on the specific synthesis methods and conditions. In some cases, these methods may lead to coatings with nanocrystallite morphologies. Small-particle-on-large-particle coating methods reported by Chen and Dahn⁴⁶ and Zhao *et al.*⁴⁷ are examples of the formation of crystalline coatings on the cathode surface. Nevertheless, this work has only focused on the crystalline polymorphs of the LiBO_2 coating rather than its amorphous ones, as the former are the simplest to study using atomistic modeling techniques and serve as a benchmark for assessing coating performance in more complex thin films and their interface with the electrodes.¹⁶ By concentrating on electron transport, this study complements our earlier investigation of ion transport and lays the groundwork for future exploration of key properties of LiBO_2 , including its thermodynamic and electrochemical stability, mechanical integrity, and interfacial characteristics. Overall, our combined DFT findings underscore the importance of lattice vacancy manipulation, particularly through techniques like thermal neutron irradiation, in advancing LiBO_2 material as a highly functional material for next-generation energy storage technologies.

Author contributions

The contributions of the authors to the current work are as follows: conceptualization: H. M. N., C.W., T. W. H., Y. X., and J. G.; methodology: H. M. N., C. Z., N. N., and C. W.; manuscript writing: H. M. N., C. W., and C. Z.; manuscript proofreading and reviewing: all authors; equal contributions: H. M. N and C. Z.

Conflicts of interest

The authors declare no conflicts of interest.

Data availability

It is officially and firmly stated that the data that support the findings of this study are available from the corresponding authors upon reasonable request. The computational models, parameters, and raw output data generated during the current study have been archived and can be shared with interested researchers to facilitate replication or further exploration of the results.

Acknowledgements

This work was funded in part by the National Science Foundation Grant No. IIP-2044726, the University of Missouri Materials Science and Engineering Institute (MUMSEI) Grant No.



CD002339, and the University of Missouri Research Reactor (MURR). The computation for this work was performed on the high-performance computing infrastructure provided by Research Support Services at the University of Missouri, Columbia, MO. DOI: <https://doi.org/10.32469/10355/97710>. The careful manuscript proofreading and valuable comments of Dr Bikash Saha (MURR) to improve the quality of this work are greatly appreciated.

References

- 1 M. M. Islam, *et al.*, The ionic conductivity in lithium-boron oxide materials and its relation to structural, electronic and defect properties: insights from theory, *J. Phys.: Condens. Matter*, 2012, **24**, 203201.
- 2 X. Jiang, *et al.*, Anomalous mechanical materials squeezing three-dimensional volume compressibility into one dimension, *Nat. Commun.*, 2020, **11**, 5593.
- 3 I. N. Ogorofnikov, *et al.*, Radiation effects and defects in lithium borate crystals, *IOP Conf. Ser.: Mater. Sci. Eng.*, 2010, **15**, 012016.
- 4 S. Zabarjad, *et al.*, Ab-initio Investigation of the Structural Stability, Electronic and Optical Properties of the LiBO₂ Compound by Using the G₀W₀ + BSE Approach, *Comput. Condens. Matter*, 2023, **34**, e00789.
- 5 S. Gao, *et al.*, Boron Doping and LiBO₂ Coating Synergistically Enhance the High-Rate Performance of LiNi_{0.6}-Co_{0.1}Mn_{0.3}O₂ Cathode Materials, *ACS Sustainable Chem. Eng.*, 2021, **9**, 5322–5333.
- 6 M. Guo, *et al.*, Excellent electrochemical properties of Ni-rich LiNi_{0.88}Co_{0.09}Al_{0.03}O₂ cathode materials co-modified with Mg-doping and LiBO₂-coating for lithium ion batteries, *New J. Chem.*, 2023, **47**, 968–2977.
- 7 B. Ramkumar, *et al.*, LiBO₂-modified LiCoO₂ as an efficient cathode with garnet framework Li_{6.75}La₃Zr_{1.75}Nb_{0.25}O₁₂ electrolyte toward building all-solid-state lithium battery for high-temperature operation, *Electrochim. Acta*, 2020, **359**, 136955.
- 8 X.-D. Zhang, *et al.*, An effective LiBO₂ coating to ameliorate the cathode/electrolyte interfacial issues of in LiNi_{0.6}-Co_{0.2}Mn_{0.2}O₂ solid-state Li batteries, *J. Power Sources*, 2019, **426**, 242–249.
- 9 P. Guan, *et al.*, Recent progress of surface coating on cathode materials for high-performance lithium-ion batteries, *J. Energy Chem.*, 2020, **43**, 220–235.
- 10 X. Tan, *et al.*, Recent progress in coatings and methods of Ni-rich LiNi_{0.8}Co_{0.1}Mn_{0.1}O₂ cathode materials: A short review, *Ceram. Int.*, 2020, **46**, 21888–21901.
- 11 G. Kaur, *et al.*, Review-Surface Coatings for Cathodes in Lithium Ion Batteries: From Crystal Structures to Electrochemical Performance, *J. Electrochem. Soc.*, 2022, **169**, 043504.
- 12 V. A. Maske and A. P. More, Conformal coatings for lithium-ion batteries: A comprehensive review, *Prog. Org. Coat.*, 2024, **188**, 108252.
- 13 S. Ghosh, *et al.*, A Review on High-Capacity and High-Voltage Cathodes for Next-Generation Lithium-ion Batteries, *J. Energy Power Technol.*, 2022, **4**, 1–59.
- 14 W. Hou, *et al.*, Recent Advances in Electrolytes for High-Voltage Cathodes of Lithium-Ion Batteries, *Trans. Tianjin Univ.*, 2023, **29**, 120–135.
- 15 B. Li, *et al.*, A Review of Solid Electrolyte Interphase (SEI) and Dendrite Formation in Lithium Batteries, *Electrochem. Energy Rev.*, 2023, **6**, 7.
- 16 S. Xu, R. M. Jacobs, H. M. Nguyen, S. Hao, M. Mahanthappa, C. Wolverton and D. Morgan, Lithium transport through lithium-ion battery cathode coatings, *J. Mater. Chem. A*, 2015, **3**, 17248–17272.
- 17 J. S. Park, *et al.*, Ultrathin Lithium-Ion Conducting Coatings for Increased Interfacial Stability in High Voltage Lithium-Ion Batteries, *Chem. Mater.*, 2014, **26**, 3128–3134.
- 18 L. Wang, *et al.*, Synergistic construction of uniform LiAlO₂ coating on LiNi_{0.8}Co_{0.1}Mn_{0.1}O₂ cathode materials by acid etching and polyvinylpyrrolidone cross-linking approaches for high-performance Li-ion batteries, *J. Alloys Compd.*, 2024, **990**, 174411.
- 19 D. Wang, *et al.*, Multifunctional Li₂O-2B₂O₃ coating for enhancing high voltage electrochemical performances and thermal stability of layered structured LiNi_{0.5}Co_{0.2}Mn_{0.3}O₂ cathode materials for lithium ion batteries, *Electrochim. Acta*, 2015, **174**, 1225–1233.
- 20 M. Du, *et al.*, Enhanced high-voltage cycling stability of Ni-rich LiNi_{0.8}Co_{0.1}Mn_{0.1}O₂ cathode coated with Li₂O-2B₂O₃, *J. Alloys Compd.*, 2019, **805**, 991–998.
- 21 S. N. Lim, *et al.*, Enhanced elevated-temperature performance of Li(Ni_{0.8}Co_{0.15}Al_{0.05})O₂ electrodes coated with Li₂O-2B₂O₃ glass, *Electrochim. Acta*, 2014, **136**, 1–9.
- 22 G. Guzman-Gonzalez, *et al.*, Borate-Based Artificial Solid-Electrolyte Interphase Enabling Stable Lithium Metal Anodes, *ACS Appl. Mater. Interfaces*, 2024, **16**, 66819–66825.
- 23 Z. Liua, *et al.*, Functional lithium borate salts and their potential application in high performance lithium batteries, *Coord. Chem.*, 2015, **292**, 56–73.
- 24 E. Hirose, *et al.*, Lithium ionic conductivities of α -LBO₂ with two-dimensional Li-Li networks and γ -LBO₂ with three-dimensional ones synthesized under high pressure, *J. Solid State Chem.*, 2019, **274**, 100–104.
- 25 M. M. Islam, *et al.*, Formation and Mobility of Li Point Defect in LiBO₂: A First-Principles Investigation, *J. Phys. Chem. C*, 2011, **115**, 12343–12349.
- 26 C. Ziemke, H. M. Nguyen, S. Amaya-Roncancio, J. Gahl, Y. Xing, T. W. Heitmann and C. Wexler, Formation of Lattice Vacancies and Their Effects on Li-ion Transport in LiBO₂ Crystals: A Comparative *ab initio* Study, *J. Mater. Chem. A*, 2025, **13**, 3146–3162.
- 27 G. S. Was, *Fundamentals of Radiation Material Science: Metals and Alloys*, Springer-Verlag, Berlin Heidelberg, 2007.
- 28 Y. M. Basalae, *et al.*, Electronic and Vibrational Properties of LiBO₂ Crystals, *Russ. Phys. J.*, 2019, **61**, 1861.



- 29 K. Schwarz and P. Blaha, Solid state calculations using WIEN2k, *Comput. Mater. Sci.*, 2003, **28**, 259–273.
- 30 P. Blaha, K. Schwarz, F. Tran, R. Laskowski, G. K. H. Madsen and L. D. Marks, Wien2k: An APW lo program for calculating the properties of solids, *J. Chem. Phys.*, 2020, **152**, 074101.
- 31 P. Blaha, K. Schwarz, P. Sorantin and S. B. Trickey, Full-potential, linearized augmented plane wave programs for crystalline systems, *Comput. Phys. Commun.*, 1990, **59**, 399–415.
- 32 J. P. Perdew, K. Burke and M. Ernzerhof, Generalized Gradient Approximation Made Simple, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 33 P. E. Blochl, O. Jepsen and O. K. Andersen, Improved tetrahedron method for Brillouin-zone integrations, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **49**, 16223.
- 34 J. Qiu, *et al.*, Effects of neutron and gamma radiation on lithium-ion batteries, *Nucl. Instrum. Methods Phys. Res., Sect. B*, 2015, **345**, 27–37.
- 35 C. Tan, *et al.*, Radiation effects on the electrode and electrolyte of a lithium-ion battery, *J. Power Sources*, 2016, **318**, 242–250.
- 36 L. Li, *et al.*, Neutron radiation on tin anodes of lithium-ion batteries, *Radiat. Eff. Defects Solids*, 2018, **173**, 1068–1074.
- 37 A. D. Pathak, *et al.*, A review on battery technology for space application, *J. Energy Storage*, 2023, **61**, 106792.
- 38 M. Takada, *et al.*, Characterization of a real-time neutron detector for boron neutron capture therapy using a thin silicon diode, *Radiat. Meas.*, 2020, **137**, 106381.
- 39 J. C. Philipps, *et al.*, Polymeric Interlayer Strengthening with Boron Neutron Capture Radiation Treatment for Laminated Glass, *Polymers*, 2023, **15**, 1672.
- 40 H. M. Nguyen, C. Ziemke, Z. Vogelsmeier, N. Narayanan, D. Stalla, B. C. Saha, J. Gahl, Y. Xing, C. Wexler and T. W. Heitmann, Is Thermal-Neutron Irradiation a Possible Approach to Defect Engineering of Advanced Materials for Their Enhanced Ionic Conductivity? Yes, It Is!, manuscript in preparation.
- 41 W. Setyawan and S. Curtarolo, High-throughput electronic band structure calculations: Challenges and tools, *Comput. Mater. Sci.*, 2010, **490**, 299–312.
- 42 A. Kokalj, XcrysDen—a new program for displaying crystal-line structures and electron densities, *J. Mol. Graphics Modell.*, 1999, **17**, 176.
- 43 D. A. Neaman, *Semiconductor Physics and Devices: Basic Principles*, McGraw-Hill, New York, 4th edn, 2012.
- 44 F. Zheng, Review on solid electrolytes for all-solid-state lithium-ion batteries, *J. Power Sources*, 2018, **389**, 1873.
- 45 R. J. D. Tilley, *Defects in Solids*, John Wiley & Sons, Inc., Hoboken, New Jersey, 2008.
- 46 Z. Chen and J. R. Dahn, Effect of a ZrO₂ Coating on the Structure and Electrochemistry of Li_xCoO₂ When Cycled to 4.5 V, *Electrochem. Solid-State Lett.*, 2002, **5**, A213–A216.
- 47 J. Zhao, G. Qu, J. C. Flake and Y. Wang, Low temperature preparation of crystalline ZrO₂ coatings for improved elevated-temperature performances of Li-ion battery cathodes, *Chem. Commun.*, 2012, **48**, 8108–8110.

