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Synthesis, bridgehead functionalization, and photoisomerization of 9,10-diboratatriptycene dianions†

Sven E. Prey,^a Jannik Gilmer,^a Samira V. Teichmann,^a Luis Čaić,^a Mischa Wenisch,^a Michael Bolte,^a Alexander Virovets,^a Hans-Wolfram Lerner,^a Felipe Fantuzzi^b and Matthias Wagner^{*a}

9,10-Diboratatriptycene salts $M_2[RB(\mu-C_6H_4)_3BR]$ ($R = H, Me$; $M^+ = Li^+, K^+, [n-Bu_4N]^+$) have been synthesized via $[4 + 2]$ cycloaddition between doubly reduced 9,10-dihydro-9,10-diboraanthracenes $M_2[DBA]$ and benzyne, generated *in situ* from C_6H_5F and C_6H_5Li or $LiN(i-Pr)_2$. $[HB(\mu-C_6H_4)_3BH]^{2-}$ reacts with CH_2Cl_2 to form quantitatively the bridgehead-derivatized $[ClB(\mu-C_6H_4)_3BCl]^{2-}$, while twofold H^- abstraction with $B(C_6F_5)_3$ in the presence of SMe_2 leads cleanly to the diadduct $(Me_2S)B(\mu-C_6H_4)_3B(SMe_2)$. Photoisomerization of $K_2[HB(\mu-C_6H_4)_3BH]$ (THF, medium-pressure Hg lamp) provides facile access to diborabenz[a]fluoranthenes, a little explored form of boron-doped polycyclic aromatic hydrocarbons. According to DFT calculations, the underlying reaction mechanism consists of three main steps: (i) photoinduced di- π -borate rearrangement, (ii) “walk reaction” of a BH unit, and (iii) boryl anion-like C–H activation.

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Introduction

The terms “aromatic hydrocarbon” and “planar π system” can be used as synonyms for most representatives of these substance classes. Triptycenes are among the few aromatic hydrocarbons that enter the third dimension, which is one reason why they have fascinated the chemical community for decades. Triptycene is built upon a $[2.2.2]$ -bicyclic system with three 1,2-phenylene blades to generate a D_{3h} -symmetric paddle-wheel scaffold.^{1,2} Due to its exceptional rigidity, the triptycene framework has been used as a backbone for transition-metal ligands^{3,4} and as a building block for molecular motors.⁵ When incorporated into emitting polymers, triptycene units can sterically block the individual polymer chains so that the materials exhibit solution-like luminescence properties in the solid state, with applications in sensor technology.⁶ The three-dimensional shape of a triptycene building block creates a free volume around the molecule (*i.e.*, between the blades of the paddle-wheel structure), which can be designed such that it confers size selectivity in sensory responses,⁷ promotes the

desired alignment in oriented polymers,⁸ or creates porous organic materials.⁹ Of more academic interest – but no less fascinating – is the excited-state reactivity of triptycene and the resulting photorearrangement products (*i.e.*, a norcaradiene and a benzo[a]fluoranthene derivative).^{10,11}

A proper means of expanding the chemical space of triptycenes as well as their application potential is heteroatom doping, *i.e.*, the replacement of selected C atoms by other p-block elements. Our group has a long-standing interest in triptycenes featuring B atoms at their bridgehead positions. One example is the $(BN)_2$ -triptycene **A** (Fig. 1). **A**-type structures are accessible by simple mixing of 9,10-dihydro-9,10-diboraanthracenes (DBAs) with 1,2-diazenes, such as pyridazine. Since the reaction is based on the essentially barrierless formation of a double B–N adduct, **A** is obtained in very good yields of $\approx 80\%$.^{12–14} An organometallic $(BN)_2$ -triptycene is the pyrazabole-bridged *ansa*-ferrocene **B** (Fig. 1).¹⁵ Recently, our group succeeded in abstracting one of its bridgehead H^- ligands and characterized the resulting free, cationic Lewis superacid $[B - H]^+$ by X-ray crystallography.¹⁶ Berionni and coworkers reported a series of boratatriptycenes **C** (Fig. 1) possessing various different atoms/groups “E” at the second bridgehead position ($E = CH, PH^+, S^+, Se^+$).^{17–20} A derivative of **C** combining $E = CH$ and $Ar = 4-t-Bu-C_6H_4$ underwent selective protodeboronation of the exocyclic C–B bond upon treatment with bistriflimidic acid ($HNTf_2$); the resulting boratriptycene Lewis acid still engages in a dynamic association–dissociation equilibrium with the weakly coordinating $[NTf_2]^-$ anion.¹⁸ This finding underscores the potential for enhancing Lewis acidity

^aInstitut für Anorganische und Analytische Chemie, Goethe-Universität Frankfurt, Max-von-Laue-Straße 7, D-60438 Frankfurt am Main, Germany. E-mail: matthias.wagner@chemie.uni-frankfurt.de

^bSchool of Chemistry and Forensic Science, University of Kent, Park Wood Rd, Canterbury CT2 7NH, UK

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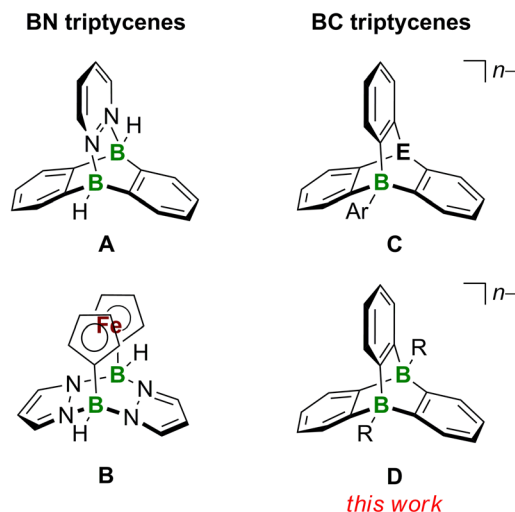


Fig. 1 Boron-containing triptycenes. Left: the DBA-pyridazine adduct **A** and the 1,1'-ferrocenediyl-bridged pyrazabole **B** as BN derivatives of triptycene. Right: Berionni's BC-triptycenes **C** featuring one B atom at a bridgehead position ($E = \text{CH}$, $n = 1$; $E = \text{PH}^+$, S^+ , Se^+ , $n = 0$; $\text{Ar} = \text{aryl}$) and the dibora(ta)triptycenes **D** presented in this work ($R = \text{H}$, Me , $n = 2$; $R = \text{SMe}_2$, $n = 0$). DBA = 9,10-dihydro-9,10-diboraanthracene.

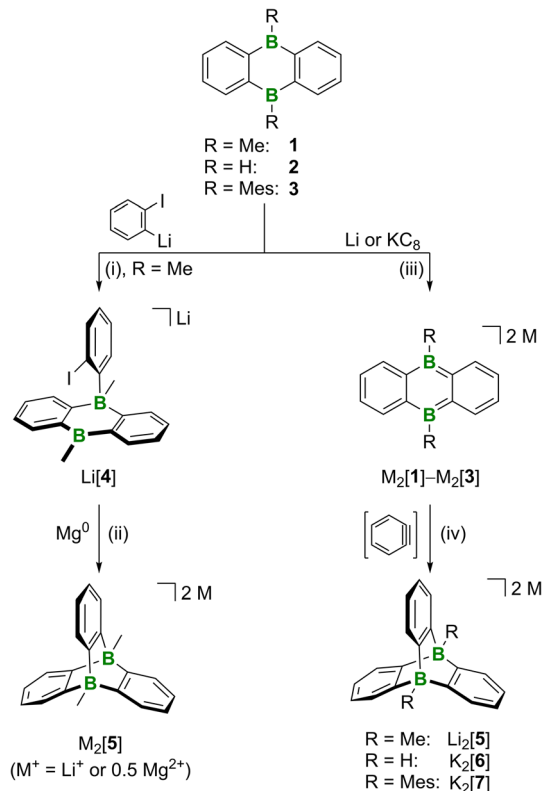
by designing a rigid organic framework in a way to impose a structural constraint on the BC_3 core that forces it out of planarity (see also the Lewis superacidity of $[\text{B} - \text{H}]^+$ mentioned above¹⁶).¹⁷

From the application examples of triptycenes, it is obvious that the degree of repulsive or attractive interactions between the paddle-wheels' blades and their immediate environments play an important role. Since these interactions will be strongly influenced by the overall charge of the triptycene scaffold, we are herein targeting dianionic diboratriptycenes **D** (Fig. 1; $R = \text{H}$, Me ; $n = 2$). We will show that H^- abstraction *in situ* releases a Lewis acid with potentially cooperating B atoms²¹ and allows convenient derivatization of both bridgehead positions. We finally show that the photorearrangement of **D** ($R = \text{H}$; $n = 2$) provides access to the little explored compound class of diborabenzo[a]fluoranthenes.²²

Results and discussion

Synthesis of 9,10-diboratriptycenes **D**

We tested two conceptually different ways to prepare our **D**-type target compounds: (a) adduct formation on neutral DBA and (b) $[4 + 2]$ cycloaddition on doubly reduced $[\text{DBA}]^{2-}$. The first approach (a) is inspired by the facile synthesis of **A** and based on the assumption that pyridazine can, in principle, be replaced by the isoelectronic benzene-1,2-diide to assemble the tricyclic scaffold of **D** *via* a twofold B–C-adduct formation. Unfortunately, the available literature on 1,2-dimetalated benzenes indicates a rather limited synthetic value of these highly reactive species.²³ We therefore resorted to a stepwise reaction protocol, starting from *in situ*-prepared *o*-iodophenyllithium and **1** (-110°C , THF/ Et_2O (1 : 1); Scheme 1).²⁴ Cyclization of the obtained $\text{B}(\text{sp}^2)\text{--B}(\text{sp}^3)$ intermediate $\text{Li}[4]$ was subsequently



Scheme 1 Synthesis of 9,10-diboratriptycenes by either sequential B–C-adduct formation (left) or $[4 + 2]$ cycloaddition of *in situ*-generated benzyne with $\text{M}_2[\text{DBA}]$ salts (right). (i) 1,2- $\text{C}_6\text{H}_4\text{I}_2$, $t\text{-BuLi}$, THF/ Et_2O (1 : 1), $-110^\circ\text{C} \rightarrow$ room temperature, overnight, 21% yield; (ii) 1,2- $\text{C}_2\text{H}_4\text{Br}_2$ (6 mol%), THF, room temperature, 3 d; (iii) THF, room temperature, 1 h, quant. conv. by NMR; (iv) $\text{C}_6\text{H}_5\text{F}$ (1–2 equiv.), base ($\text{LiN}(\text{i-Pr})_2$ or $\text{C}_6\text{H}_5\text{Li}$), THF, $-78^\circ\text{C} \rightarrow$ room temperature, overnight to 5 d, $\text{Li}_2[5]$: 50–75% yield, $\text{K}_2[6]$: 41% yield, $\text{K}_2[7]$: few crystals.

accomplished by stirring with Mg^0 (room temperature, THF). Since the NMR-spectroscopic investigation of the crude product still showed the formation of some unwanted side products that later proved difficult to remove, the B–C-adduct approach was eventually abandoned in favor of the Umpolung strategy (b): given that doubly reduced DBAs easily undergo $[4 + 2]$ -cycloaddition reactions with (hetero)olefins or internal alkynes,²⁵ the reaction of DBA salts $\text{M}_2[1]\text{--M}_2[3]$ with *in situ*-generated benzyne should likewise generate diboratriptycenes $\text{M}_2[5]\text{--M}_2[7]$ (Scheme 1). For an initial proof-of-principle experiment, we selected $\text{Li}_2[1]$, because we have had good experience with cycloadditions on this compound in the past.²⁵ $\text{LiN}(\text{i-Pr})_2$ in THF was added with stirring at -78°C to a mixture of freshly prepared $\text{Li}_2[1]$ (ref. 26 and 27) and $\text{C}_6\text{H}_5\text{F}$ in THF (Scheme 1). After warming to room temperature overnight, the reaction mixture had lost the characteristic red color of $\text{Li}_2[1]$ and turned to pale yellow. ^{11}B NMR spectroscopy showed the selective formation of one new species; its chemical shift value of $\delta(^{11}\text{B}) = -13.2$ ppm testified to the presence of tetra-coordinated B atoms.²⁸ An ^1H NMR spectrum, recorded on the crude product after a solvent change to THF- d_8 , indicated an essentially quantitative conversion: a signal at 0.56 ppm (6H)

and two resonances at 7.24 (br, 6H) and 6.44 ppm (m, 6H) are assignable to two chemically equivalent CH₃ groups and three chemically equivalent C₆H₄ rings, respectively. Also the ¹³C{¹H} NMR spectra supported the proposed D_{3h}-symmetric molecular structure of [5]²⁻ in solution. Purification of Li₂[5] through crystallization from THF gave yields of 50–75% (attempts at X-ray analysis were not successful). Next, we moved on from DBA 1 to DBA 2 to also get a hold of the parent diboratriptycene. K₂[6] was prepared in a similar manner as Li₂[5] and isolated in 41% yield ($\delta(^{11}\text{B}) = -8.3$ ppm, d, $^1J(\text{B},\text{H}) = 87$ Hz); the BH resonance was found at 3.36 ppm in the ¹H{¹¹B} NMR spectrum.²⁹ To stress the steric limits of the [4 + 2] cycloaddition, we finally treated dimesitylated K₂[3] with benzyne and still obtained K₂[7], albeit in negligible yield of only a few crystals.²⁴

Crystals of [K(dme)][K(dme)₂][6] suitable for X-ray diffraction (XRD) were obtained from a gas-phase diffusion experiment (*n*-hexane/1,2-dimethoxyethane (DME); Fig. 2). In [K(dme)][K(dme)₂][6], the diboratriptycene dianion [6]²⁻ acts as chelating ligand toward both K⁺ ions (*via* the π -electron clouds of its 1,2-phenylene rings). The coordination sphere of K(1)⁺ is completed by two dme ligands; K(2)⁺ carries only one dme ligand and is further bonded to two phenylene rings of an adjacent [6]²⁻ moiety, thereby creating a coordination polymer in which all three pockets of the paddle-wheel anion are occupied by K⁺ ions (Fig. S61†). The six B–C bond lengths of C₁-symmetric (in the crystal lattice) [6]²⁻ range from 1.620(3) to 1.643(3) Å (avg. 1.631 Å) and are comparable to those in the structurally unconstrained, acyclic analogue [HBPh₃][−] (1.628(6) to 1.633(6) Å, avg. 1.631 Å).³⁰ Of the three dihedral angles 107.9°, 124.8°, and 127.3° between phenylene rings, the smallest is associated with the pocket hosting the doubly η^6 -coordinated K(1)⁺.

Activation of the B–H bonds of the 9,10-diboratriptycene [6]²⁻

Having the diboratriptycene salt [K(dme)][K(dme)₂][6] in hand, we next addressed the question of how to activate its B–H bonds to derivatize the bridgehead positions and/or generate the free Lewis acid *in situ*. The neutral ditopic Lewis acid **S1** (Scheme 2) should be among the strongest neutral B-based

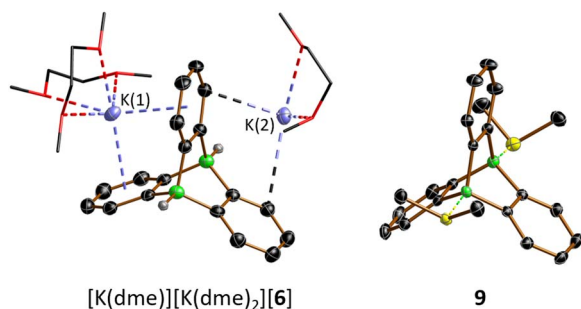
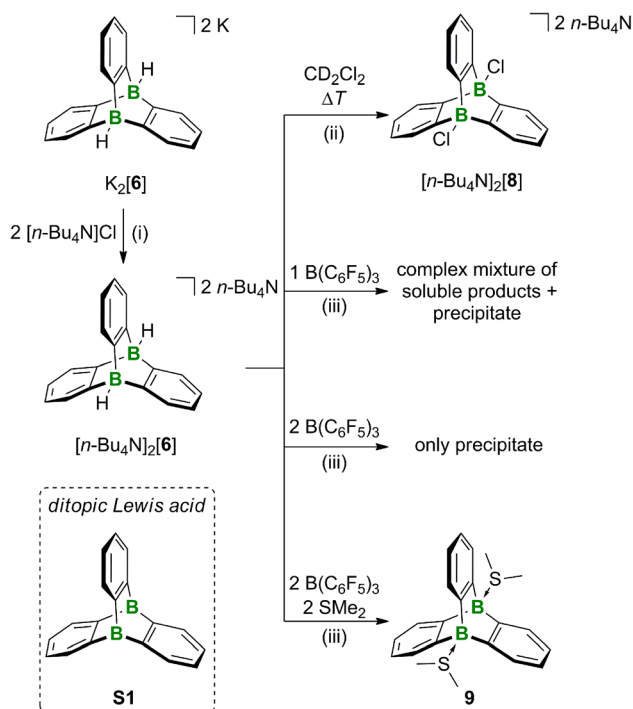


Fig. 2 Solid-state structures of [K(dme)][K(dme)₂][6] and **9** with thermal ellipsoids shown at the 50% probability level; C-bonded H atoms and a non-coordinating CH₂Cl₂ molecule in the crystal lattice of **9** are omitted for clarity, dme ligands are depicted as wireframes. H: gray, B: green, C: black, O: red, S: yellow, K: light blue.



Scheme 2 Reactivity of 9,10-diboratriptycene [6]²⁻. (i) (1) THF, room temperature, 1 h; (2) extraction into CH₂Cl₂, 93% yield; (ii) 50 °C, 18 h, 89% yield; (iii) CD₂Cl₂, room temperature, 5 min, **9**: quant. conv. by NMR.

acids, as judged by its computed F[−]-ion affinity FIA_{solv}(**S1**) of 384 kJ mol^{−1} (*cf.* FIA_{solv}(B(C₆F₅)₃) = 315 kJ mol^{−1}; solv = CH₂Cl₂; for FIAs of further B₂-triptycenes and a more detailed discussion see the ESI†).¹⁶ As the dme ligands present in the complex cations could potentially become B-coordinating and thereby interfere with the planned reactivity studies, we first performed a cation exchange with [n-Bu₄N]Cl in THF. [n-Bu₄N]₂[6] precipitates together with KCl from the reaction mixture and was separated from the latter by extraction into CH₂Cl₂ (93% yield). Care must be taken to remove the solvent from the product in time after workup to avoid B–H/B–Cl exchange.³¹ On the other hand, the reaction between [n-Bu₄N]₂[6] and CH₂Cl₂ can be used for the straightforward synthesis of bridgehead-chlorinated [n-Bu₄N]₂[8] (14 d at room temperature or 18 h at 50 °C; 89% yield; Scheme 2). In an attempt to release the free monotopic Lewis acid, [n-Bu₄N]₂[6] was treated with 1 equiv. of B(C₆F₅)₃ in CD₂Cl₂. NMR spectroscopy showed that B(C₆F₅)₃ had been quantitatively transformed to [n-Bu₄N][HB(C₆F₅)₃], even though some unconsumed starting material [n-Bu₄N]₂[6] remained in the mixture. Several non-identified species and small quantities of a precipitate had formed. When the amount of B(C₆F₅)₃ added was increased to 2 equiv., the H[−] scavenger was still fully converted, phenylene resonances were no longer detectable in the ¹H NMR spectrum, and substantial amounts of precipitate were found in the NMR tube. A similar result was obtained when [Ph₃C][B(C₆F₅)₄] was used instead of B(C₆F₅)₃. The precipitate could not be re-dissolved by addition of SME₂. The situation changed when the double H[−] abstraction was carried

out with $\text{B}(\text{C}_6\text{F}_5)_3$ in the presence of SMe_2 , which led to the clean formation of the diboratriptycene- SMe_2 diadduct **9** (Scheme 2). Based on the computed FIAs of relevant B_2 -triptycenes, we propose that the formation of **9** proceeds in a stepwise manner without intermediate formation of **S1**.²⁴ X-ray quality crystals of $9 \cdot \text{CH}_2\text{Cl}_2$ were grown from CH_2Cl_2 (Fig. 2).

The ^{11}B NMR spectrum of **9** is characterized by a signal at -1.4 ppm, in the region of tetracoordinated B nuclei.²⁸ The $^1\text{H}/^{13}\text{C}$ resonances of the SMe_2 ligands appear at $3.07/21.1$ ppm (CD_2Cl_2), significantly downfield-shifted from those of free SMe_2 ($2.09/18.4$ ppm) and even $\text{Me}_2\text{S}-\text{B}(\text{C}_6\text{F}_5)_3$ ($2.17/20.8$ ppm;³² CDCl_3), thus indicating a strong S-B-adduct bond. This conclusion is further confirmed by the short S-B-bond lengths of **9** ($1.9710(19)$, $1.9775(18)$ Å; cf. $\text{Me}_2\text{S}-\text{B}(\text{C}_6\text{F}_5)_3$: S-B = $2.091(5)$ Å (ref. 33)).

To summarize, we assume that after H^- abstraction on $[\mathbf{6}]^{2-}$, the corresponding free Lewis acid **S1** is not stable but rather forms the insoluble solid by rearrangement or ring-opening polymerization. Although the structure of the solid remains to be confirmed, a plausible candidate is a polymer composed of 1,2-phenylene-bridged DBAs (Fig. S15†). The formation of a precipitate is suppressed if SMe_2 is present from the beginning to stabilize the diboratriptycene.

Photoisomerization of the 9,10-diboratriptycene $[\mathbf{6}]^{2-}$

In contrast to the thermal rearrangement of free neutral diboratriptycene, photoisomerization of the diboratriptycene salt $\text{K}_2[\mathbf{6}]$ results in the well-defined, soluble product $\text{K}_2[\mathbf{10}]$ (Scheme 3). $\text{K}_2[\mathbf{10}]$ can be viewed as a molecule in which a DBA moiety shares one B atom and one phenylene ring with a 9H-9-boraffluorene unit.

Irradiation of a THF solution of $\text{K}_2[\mathbf{6}]$ with a medium-pressure Hg lamp in a quartz vessel for 1 h led to a color change from colorless to yellow.³⁴ NMR spectroscopy on a THF- d_8 solution revealed the highly selective formation of $\text{K}_2[\mathbf{10}]$. Its ^{11}B NMR spectrum shows a doublet ($\delta(^{11}\text{B}) = -13.7$ ppm, $^1J(\text{B},\text{H}) = 73$ Hz) and a pseudo-triplet ($\delta(^{11}\text{B}) = -16.7$ ppm, $^1J(\text{B},\text{H}) = 78$ Hz), assignable to a BH- and a BH_2 group, respectively. The corresponding BH_n resonances were found in the $^1\text{H}\{^{11}\text{B}\}$ NMR spectrum at 2.65 (BH) and $2.52/2.40$ ppm (BH_2). The $^{13}\text{C}\{^1\text{H}\}$ NMR spectrum supports the proposed C_{1v} -symmetric molecular structure, since we observe 18 signals belonging to three different six-membered rings (according to $^1\text{H}-^{13}\text{C}$ -HSQC and -HMBC experiments). Of those signals, five are severely broadened due to $^1J(\text{B},\text{C})$ coupling and the

quadrupolar effect of the $^{10/11}\text{B}$ nuclei and therefore correspond to the five required B-bonded C atoms; deshielded, quaternary ^{13}C atoms giving rise to NMR resonances at $151.1/144.3$ ppm point toward the presence of a biphenyl fragment in the molecule. Moreover, the DFT-computed ^{13}C shift values of $[\mathbf{10}]^{2-}$ agree well with the experimentally observed ones for $\text{K}_2[\mathbf{10}]$ (avg. deviation: 2.0 ppm).²⁴ In order to gain more information on the mechanism underlying the photoisomerization of $[\mathbf{6}]^{2-}$ to $[\mathbf{10}]^{2-}$, we performed corresponding quantum-chemical calculations.

Assessment of the photoisomerization of $\text{K}_2[\mathbf{6}]$ by theory

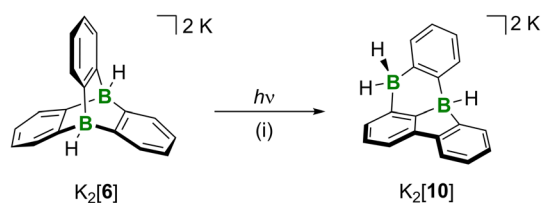
All energies were calculated at the SMD(THF)³⁵/ ω B97XD³⁶/6-311++G** level of theory, with optimized structures at SMD(THF)/ ω B97XD/6-31+G**. Vertical excitation energies were calculated with the respective unrestricted functional (T_1 states) or by TD-DFT calculations (S_1 states). The calculations were performed on the bare dianionic species (Fig. 3a).³⁷

We took the known photoisomerization pathway of carbaceous triptycene **6**^C and the di- π -borate rearrangement of the tetraarylborate $\text{K}[\mathbf{11}]$ as blueprints for our mechanistic studies (Fig. 3b and c).^{10,11,38} The first reaction step $[\mathbf{6}]^{2-} \rightarrow [\mathbf{Int1}]^{2-}$ involves C-C coupling between two aryl rings with simultaneous extrusion of one BH unit. In line with experiment, the associated energy barrier of $62.0 \text{ kcal mol}^{-1}$ is prohibitively high for a reaction under ambient conditions (in the S_0 states). To account for the actually applied photochemical protocol, this crucial step was investigated in detail also for the excited T_1 and S_1 states of the molecules involved.

First, we verified the basic premise, namely that the Hg lamp employed (emission range: $200\text{--}600 \text{ nm}$) does indeed provide the wavelength required to overcome the computed $\text{S}_0 \rightarrow \text{S}_1$ energy gap of $[\mathbf{6}]^{2-}$ ($\Delta E_{\text{S}_0 \rightarrow \text{S}_1} = 5.068 \text{ eV}$, corresponding to $\lambda_{\text{exc}} = 244.7 \text{ nm}$; $f_{\text{osc}} = 0.1010$).

We scanned the T_1 and S_1 energy surface topographies along the reaction coordinate between $[\mathbf{6}]^{2-}$ and $[\mathbf{Int1}]^{2-}$ based on ground-state structures obtained from the $[\mathbf{TS1}]^{2-}$ -IRC scan (Fig. 4; IRC = intrinsic reaction coordinate): the energy curves for both excited states have minima located downstream of the $[\mathbf{TS1}]^{2-}$ structure along the reaction coordinate. It is reasonable that – regardless of a potential intersystem crossing – both these minima may provide access back to the ground state.³⁹ From that on, a barrierless relaxation leads to $[\mathbf{Int1}]^{2-}$ at $36.9 \text{ kcal mol}^{-1}$ relative to $[\mathbf{6}]^{2-}$. We note that boratanorcaradiene substructures, such as the one present in the key intermediate $[\mathbf{Int1}]^{2-}$, have been characterized by XRD in the form of $\text{K}[\mathbf{12}]$ and $\text{Na}[\mathbf{14}]$ (Fig. 3c and d).^{38,40}

All barriers of the following “walk reaction” on the S_0 surface are reachable at ambient conditions:⁴¹ $[\mathbf{Int1}]^{2-}$ undergoes a $[1,5]$ sigmatropic shift of the boryl group to form $[\mathbf{Int2}]^{2-}$, which is $2.7 \text{ kcal mol}^{-1}$ more stable than $[\mathbf{Int1}]^{2-}$. The corresponding barrier $[\mathbf{Int1}]^{2-} \rightarrow [\mathbf{TS2}]^{2-}$ amounts to $\Delta G^\ddagger = 17.6 \text{ kcal mol}^{-1}$. A second, slightly endergonic $[1,5]$ sigmatropic shift *via* $[\mathbf{TS3}]^{2-}$ gives $[\mathbf{Int3}]^{2-}$ ($\Delta G^\ddagger = 19.7 \text{ kcal mol}^{-1}$, $\Delta G = 7.9 \text{ kcal mol}^{-1}$). In the last step, C-H activation from $[\mathbf{Int3}]^{2-}$ leads to $[\mathbf{10}]^{2-}$.⁴² The corresponding $[\mathbf{TS4}]^{2-}$ lies $12.7 \text{ kcal mol}^{-1}$ above $[\mathbf{Int3}]^{2-}$, so that the free-energy span to be covered from $[\mathbf{Int2}]^{2-}$ to $[\mathbf{TS4}]^{2-}$



Scheme 3 Photoisomerization of $\text{K}_2[\mathbf{6}]$ to the 8,12b-dihydro-8,12b-diborabenzofluoranthene derivative $\text{K}_2[\mathbf{10}]$. (i) Medium-pressure Hg lamp, THF, room temperature, 1 h, quant. conv. by NMR.



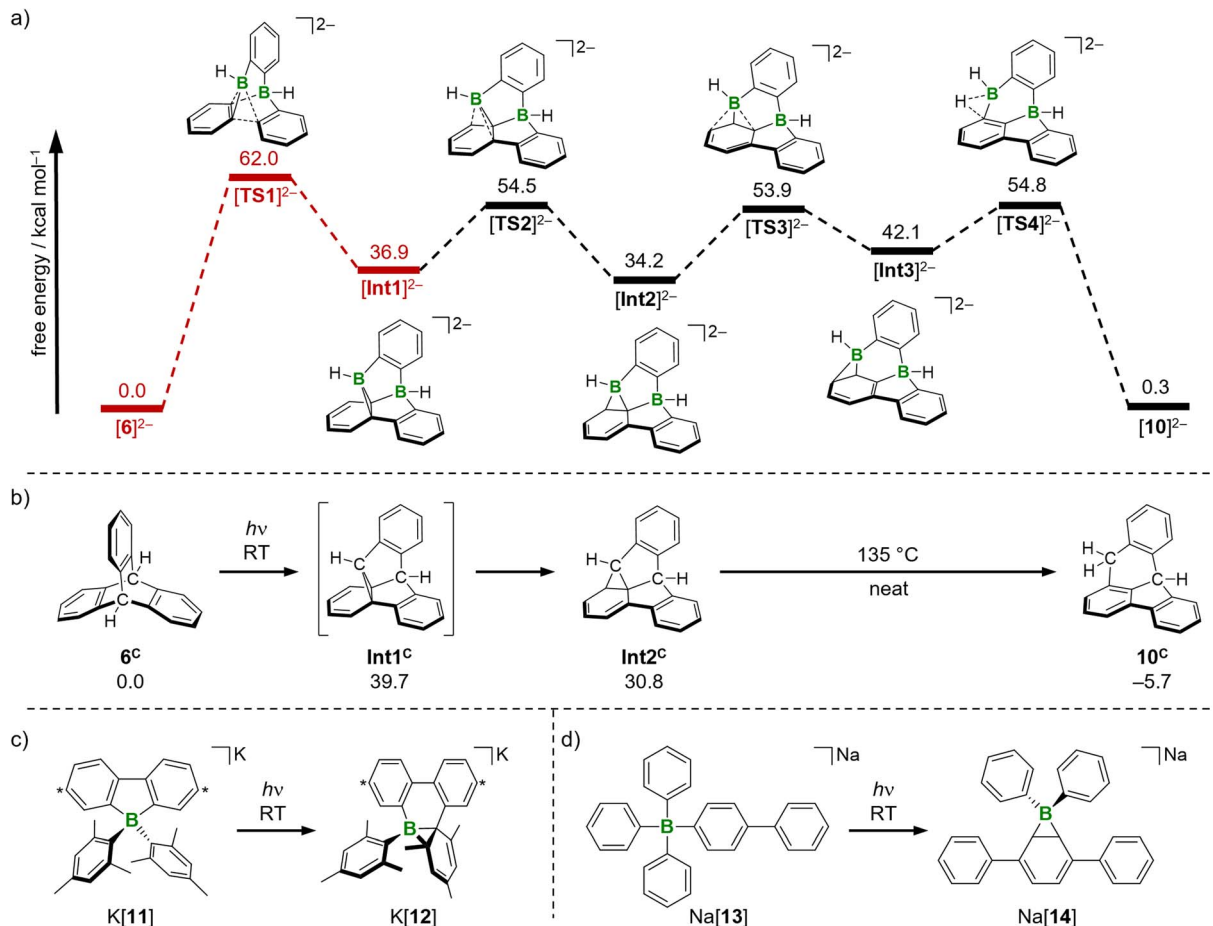


Fig. 3 (a) Computed mechanism of the photoisomerization $[6]^{2-} \rightarrow [10]^{2-}$ and free energies of intermediates and transition states in their S_0 states. The energetically unfavorable and thus photochemically conducted step $[6]^{2-} \rightarrow [Int1]^{2-}$ is marked in red. (b) Literature-reported intermediates of the photoisomerization of triptycene 6^C and their calculated free energies (this work). (c) and (d) Known boratanorcaradienes $K[12]$ and $Na[14]$ obtained from photoreactions of tetraarylborates $K[11]$ and $Na[13]$, respectively. Carbon atoms marked with asterisks bear *t*-Bu substituents.

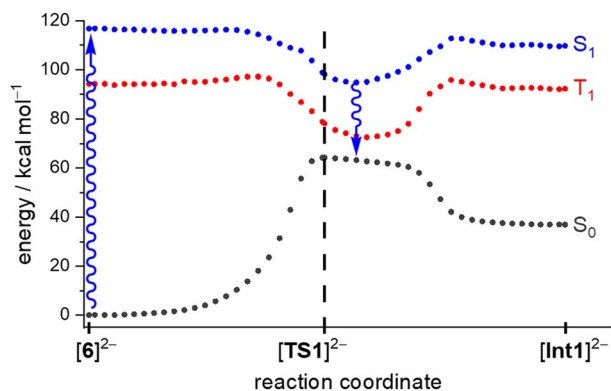


Fig. 4 Computed energy profiles of the photoinduced reaction $[6]^{2-} \rightarrow [Int1]^{2-}$ in the S_0 (gray), T_1 (red), and S_1 (blue) states. The dashed black line marks the position of $[TS1]^{2-}$ in the ground state; the blue arrows mark the electronic excitation and relaxation (the overall conclusion would remain the same if the relaxation occurred after a putative intersystem crossing from T_1).

is 20.6 kcal mol⁻¹.⁴³ The final product $[10]^{2-}$ is essentially iso-energetic to the starting material $[6]^{2-}$. Nevertheless, we consider a reverse reaction unlikely for the following reasons: (i) the energy barrier for the step $[Int1]^{2-} \rightarrow [TS1]^{2-}$ in the reverse reaction is significantly higher (25.1 kcal mol⁻¹; Fig. 3) than any barrier downstream $[Int1]^{2-}$ for the forward reaction. (ii) Assuming that the barrier $[Int1]^{2-} \rightarrow [TS1]^{2-}$ could be overcome photochemically, the system would again be in the state described in Fig. 4 and should therefore preferentially fall back to $[Int1]^{2-}$.

Conclusions

We achieved the synthesis of the parent dianionic 9,10-diboratatriptycene $[6]^{2-}$ and its bridgehead derivatization. In the salt $[K(dme)][K(dme)_2][6]$, $[6]^{2-}$ acts as a chelating ligand toward both K^+ ions, suggesting a possible future use as a tris-benzene ligand for mono- to trinuclear transition metal complexes.³ Abstraction of a bridgehead H^- substituent releases the free Lewis acid, which is sufficiently long-lived to be trapped by SME_2 ; this transformation can be carried out at both

bridgehead positions and eventually leads to the neutral boratriptycene-diadduct **9**. By using ditopic Lewis bases, this derivatization chemistry should provide an entry point to polymeric dibora(tri)ptycenes for materials science. The facile photoisomerization of $K_2[6]$ furnishes the diborabenz[a]-fluoranthene derivative $K_2[10]$, which combines the two prominent structural motifs of 9,10-dihydro-9,10-diboranthracene and 9H-9-borafluorene and belongs to a virtually unexplored class of B-doped polycyclic aromatic hydrocarbons.

Data availability

The datasets supporting this article have been uploaded as part of the ESI†

Author contributions

S. E. P. performed the experimental studies and characterized all new compounds. S. E. P. and J. G. performed the quantum-chemical calculations. S. V. T., L. C., and M. We. assisted with the synthesis of compounds **9**, $[n\text{-Bu}_4\text{N}]_2[8]$, and $\text{Li}_2[5]$, respectively. M. B. performed the X-ray crystal structure analyses of the compounds $[\text{K}_2(\text{dme})_3][3]$, $[\text{Li}(\text{thf})_2][4]$, $[\text{K}(\text{dme})][\text{K}(\text{dme})_2][6]$, and $[\text{K}(\text{dme})][\text{K}(\text{dme})_2][7]$. A. V. performed the X-ray crystal structure analysis of $9 \cdot \text{CH}_2\text{Cl}_2$. H.-W. L., F. F., and M. Wa. supervised the project. The manuscript was written by S. E. P., J. G., and M. Wa. and edited by all co-authors.

Conflicts of interest

There are no conflicts to declare.

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