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Catalytic decomposition of N₂O over Rh/Zn–Al₂O₃ catalysts†

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Zn–Al₂O₃ supports were prepared by impregnating commercial γ -Al₂O₃ powders with different amounts of Zn(NO₃)₂, followed by calcination in air at 500 or 800 °C. Rh/Zn–Al₂O₃ catalysts were then prepared by impregnating Zn–Al₂O₃ supports with Rh(NO₃)₃ followed by calcination in air at 500 °C. The catalysts and/or supports were characterized by ICP-OES, XRD, N₂ adsorption, Raman spectroscopy, TEM-EDX, XPS, CO₂-TPD, H₂-TPR, and O₂-TPD, and the catalytic performance of supported Rh catalysts in N₂O decomposition was tested. It is concluded that the support can be described as ZnO/Al₂O₃ (ZnO supported on Al₂O₃) when calcining Zn(NO₃)₂/Al₂O₃ at 500 °C, whereas ZnAl₂O₄ spinel forms on the Al₂O₃ surface at 800 °C. Rh/Zn–Al₂O₃ catalysts are much more active than Rh/Al₂O₃ and Rh/ZnO. The best catalyst (Rh/Zn–Al₂O₃-800 with 1 wt% Rh and 1 wt% Zn) has the smallest Rh₂O₃ particle size and can desorb O₂ at lower temperature than other catalysts. Both factors may be important for achieving high activity in N₂O decomposition.

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Introduction

The negative environmental impact of N₂O on global warming and ozone layer depletion has raised much concern. The concentration of N₂O in the atmosphere has been increasing at an annual rate of 0.2–0.3% since the industrial revolution.¹ Anthropogenic N₂O emission comes from several chemical processes (e.g., nitric acid and adipic acid production) and fossil fuel/biomass burning. N₂O can be decomposed into N₂ and O₂ over supported noble metal catalysts, bare or supported metal oxides, or zeolite-based catalysts.^{2–5} Rh-Based catalysts often exhibit high activity at relatively low reaction temperatures. Rh species can exist in the form of metallic Rh or Rh₂O₃, depending on whether the catalysts are reduced or not. Typical supports for loading Rh species include bare metal oxides,^{6–14} mixed/composite metal oxides,^{15–21} zeolites,^{22–24} and metal phosphates/hydroxyapatite.^{25–27}

Al₂O₃ is both a support and a catalyst widely used in industry. Rh/Al₂O₃ shows moderate catalytic activity in N₂O decomposition.^{8,10,11,25} Attempts have been made to improve Al₂O₃-based Rh catalysts for N₂O decomposition. For example, Haber *et al.* reported that the presence of alkali metal additives on Rh/Al₂O₃ can

lead to more active catalysts due to the improved dispersion of Rh species.²⁸ Parres-Escápez *et al.* found that Sr can promote the activity of Rh/Al₂O₃ due to the improved dispersion and reducibility of Rh species.²⁹ Zhao *et al.* reported that Rh/SiO₂-Al₂O₃ shows high activity, because oxygen desorption property is improved and Rh⁰ species is stabilized.³⁰ Kim and co-workers reported that Rh/Ce–Al₂O₃ is more active than Rh/Al₂O₃, due to increased surface area and improved reducibility of Rh species.³¹

It has been reported that spinel phase has a strong interaction with noble metals, resulting in smaller size and better stability of supported noble metal particles,^{32,33} which is beneficial for N₂O decomposition.^{22,25,28} High temperature calcination of Al₂O₃-supported metal nitrates is a convenient method to form spinel phase.^{34–36} However, to the best of our knowledge, there has been no work dealing with N₂O decomposition over Rh/M–Al₂O₃ catalysts with spinel phase.

Herein, we prepared Zn–Al₂O₃ supports by impregnating Zn(NO₃)₂ on commercial γ -Al₂O₃ powders followed by calcination. Rh/Zn–Al₂O₃ catalysts were synthesized by impregnating Rh(NO₃)₃ onto Zn–Al₂O₃ supports. The catalyst prepared under optimal conditions was found to be much more active than Rh/Al₂O₃ and Rh/ZnO in N₂O decomposition. The catalysts were characterized in detail, and reasons for the high activity of Rh/Zn–Al₂O₃ were elucidated.

Experimental section

Preparation

A calculated amount of Zn(NO₃)₂·6H₂O was dissolved in 20 mL deionized water in an agate mortar. Then, 1.98 g γ -Al₂O₃ powder (specific surface area = 110 m² g^{−1}) was added and the slurry

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was mixed sufficiently using an agate pestle, followed by drying under an infrared lamp. The $\text{Zn}(\text{NO}_3)_2/\text{Al}_2\text{O}_3$ precursors were calcined in a muffle furnace at a certain temperature (500, 600, 700, 800, or 900 °C) for 4 h under flowing air. The obtained supports are referred to as $x\% \text{ Zn-Al}_2\text{O}_3\text{-}y$, where $x\%$ represents the wt% Zn in the catalysts and y the calcination temperature in °C. $\text{Zn-Al}_2\text{O}_3\text{-}y$ in the text usually has 1 wt% Zn, unless otherwise specified.

For comparison, commercial $\gamma\text{-Al}_2\text{O}_3$ was also calcined at 800 °C for 4 h under flowing air. ZnO was prepared by precipitation. 50 mL ammonia solution (2.8 mol L^{-1}) was added dropwise to 100 mL $\text{Zn}(\text{NO}_3)_2$ solution (0.6 mol L^{-1}) under stirring. The precipitates were isolated by filtration, thoroughly washed by deionized water till pH value of filter liquor reached 7, dried at 100 °C overnight, and calcined at 800 °C for 4 h under flowing air. These two samples are referred to as $\text{Al}_2\text{O}_3\text{-800}$ and ZnO-800 , respectively. An additional ZnO sample was prepared by calcining $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ at 800 °C for 4 h under flowing air.

Rhodium was loaded onto supports by impregnation. 1.98 g support was mixed with 10 mL $\text{Rh}(\text{NO}_3)_3$ solution (2 mg mL^{-1}) in an agate mortar and dried under an infrared lamp (theoretical Rh content is 1 wt%). The obtained powders were calcined at 500 °C for 4 h.

Characterization

XRD patterns were recorded on a MSAL XD2 X-ray diffractometer using $\text{CuK}\alpha$ radiation at a scanning speed of 4° min^{-1} , with voltage of 40 kV and current of 30 mA. BET surface areas were measured on a Micromeritics Tristar 3000 instrument. The samples were treated at 300 °C in vacuum for 3 h, followed by N_2 adsorption at -196°C . ICP-OES was measured on a PerkinElmer OPTIMA 2100 DV optical emission spectrometer. 0.1 g sample was dissolved in a mixture of 3 mL HNO_3 , 9 mL HCl , 1 mL HClO_4 , 0.5 mL H_2O_2 , and 3 mL HF reagents, followed by heating at 150 °C for 2–3 h. After that, the sample was mixed with 1 mL HNO_3 , 3 mL HCl , 0.5 mL HClO_4 , and 1 mL HF again, and the mixture was transferred to Teflon autoclave, heated at 180 °C for 4 h, cooled to the ambient temperature, and then diluted with distilled water for analysis.

Raman spectra were recorded on a HORIBA Jobin Yvon XploRA spectrometer. In order to avoid fluorescence of ZnO, wavelength of exciting light was selected as 532 nm. TEM data were obtained by an FEI Tecnai $\text{G}^2 \text{ F20 S-TWIN}$ with an EDX instrument. All samples were measured under an accelerating voltage of 200 kV. XPS spectra were recorded on a Shimadzu/Kratos AXIS Ultra DLD spectrometer with $\text{AlK}\alpha$ radiation as the excitation source. The C1s line (284.6 eV) was used as the reference to calibrate the binding energy.

Rh dispersion was determined by CO chemisorption on a Micromeritics AutoChem II instrument, based on 1 : 1 stoichiometry (CO/Rh). 0.15 g Rh catalyst (40–60 mesh) was pretreated in He flow at 400 °C for 1 h, and cooled to 30 °C. Then, a pulse of 5% CO-He mixture was repeatedly injected into the reactor *via* a six-way valve until the CO signals from the thermal conductivity detector remained constant. He was used as the carrier gas. The volume of CO chemisorbed was determined by summing the fractions of CO consumed in each pulse.

CO_2 -TPD experiments were conducted on a Micromeritics AutoChem II instrument. 0.2 g sample (40–60 mesh) was pretreated in He flow at 500 °C for 1 h, and cooled to 80 °C. The flow was switched to 5% CO_2/He (30 mL min^{-1}) and kept for 1 h, and then swept by He (30 mL min^{-1}) for 1.5 h. Finally, the sample was heated in He (30 mL min^{-1}) to 600 °C at a rate of $10^\circ \text{C min}^{-1}$.

H_2 -TPR experiments were conducted on a FINESORB-3010 instrument, equipped with a mass spectrometer (OmniStar™). 0.15 g catalyst (40–60 mesh) was pretreated in He flow at 400 °C for 1 h, cooled down to room temperature, and further cooled in an ice-water bath. Then, the catalyst was exposed in 4% $\text{H}_2\text{-He}$ mixture (30 mL min^{-1}) at 0 °C for 1 h, and heated by a furnace to 600 °C at a rate of $10^\circ \text{C min}^{-1}$. The profile was recorded through the channel of $m/z = 2$.

O_2 -TPD experiments were conducted on a FINESORB-3010 instrument. 0.2 g catalyst (40–60 mesh) was pretreated in He flow at 500 °C for 1 h, and cooled to 50 °C. The catalyst was exposed to O_2 (10 mL min^{-1}) at 50 °C for 1 h, swept by He (30 mL min^{-1}) for 3 h, and heated in He (30 mL min^{-1}) to 580 °C at a rate of $10^\circ \text{C min}^{-1}$.

Catalytic tests

Catalytic activity of N_2O decomposition was tested in a fixed-bed flow microreactor. 0.5 g catalyst (40–60 mesh) was packed in a U-shaped glass tube (7 mm inner diameter) sealed by quartz wool. A gas stream of 0.5% N_2O (balanced by He) flowed through the catalyst at a rate of 60 mL min^{-1} . The catalyst (first kept at near room temperature) was exposed to the gas stream for 1 h during which the existing stream was periodically analyzed by a gas chromatograph (GC, Agilent 7890A) equipped with columns (Sepuco 6ftQ and Sepuco Q&5A, column temperature: 80 °C) that can separate N_2O , O_2 , and N_2 . The reaction temperature was then raised using a furnace and kept at various elevated temperature for 0.5 h in each temperature step. The exhaust was again periodically analyzed by the GC, and the conversion of N_2O was calculated according to $X = ([\text{N}_2\text{O}]_{\text{in}} - [\text{N}_2\text{O}]_{\text{out}})/[\text{N}_2\text{O}]_{\text{in}}$, where $[\text{N}_2\text{O}]_{\text{in}}$ refers to the N_2O concentration or peak area at room temperature, and $[\text{N}_2\text{O}]_{\text{out}}$ refers to the N_2O concentration or peak area at an elevated temperature.

Results and discussion

Structural and physical properties

$\text{Rh}/\text{Zn-Al}_2\text{O}_3\text{-500}$ and $\text{Rh}/\text{Zn-Al}_2\text{O}_3\text{-800}$ catalysts were prepared by using different supports prepared by calcining $\text{Zn}(\text{NO}_3)_2/\text{Al}_2\text{O}_3$ in air at 500 or 800 °C. The Zn and Rh contents are both fixed to be 1 wt%. In addition, $\text{Rh}/\text{Al}_2\text{O}_3\text{-800}$ and $\text{Rh}/\text{ZnO-800}$ were also prepared for comparison. The Rh contents of these four catalysts are determined by ICP-OES as 1.1 wt%, 1.1 wt%, 1.0 wt%, and 1.0 wt%, respectively. The Zn contents of $\text{Rh}/\text{Zn-Al}_2\text{O}_3\text{-500}$ and $\text{Rh}/\text{Zn-Al}_2\text{O}_3\text{-800}$ are determined as 1.1 wt% and 1.1 wt%, respectively, in accordance with the theoretical value.

Fig. 1 shows the XRD patterns of supported Rh catalysts. $\text{Rh}/\text{Zn-Al}_2\text{O}_3\text{-500}$, $\text{Rh}/\text{Zn-Al}_2\text{O}_3\text{-800}$, and $\text{Rh}/\text{Al}_2\text{O}_3\text{-800}$ exhibit identical diffraction peaks at *ca.* 33, 37, 40, 46, 60, and 67° , in accordance with the diffraction peaks of $\gamma\text{-Al}_2\text{O}_3$ (PDF#47-1308).



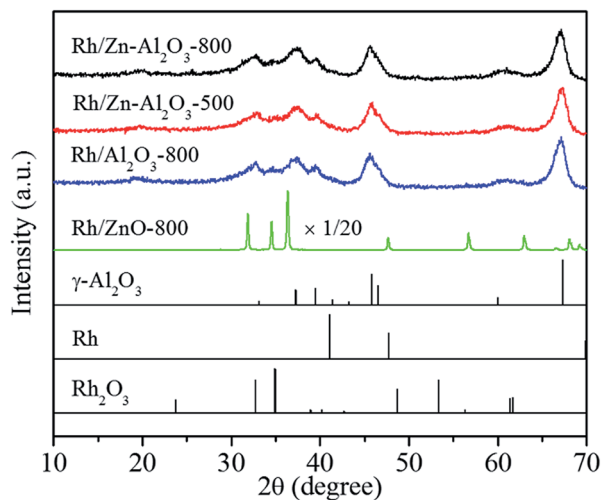


Fig. 1 XRD patterns of the catalysts and standard patterns of γ - Al_2O_3 (PDF#47-1308), Rh (PDF#05-0685), and Rh_2O_3 (PDF#41-0541).

In addition, no Rh or Rh_2O_3 can be detected, due to the low Rh content (*ca.* 1 wt%) and high dispersion, as demonstrated by TEM and Rh dispersion data later. Rh/ZnO-800 shows very sharp peaks corresponding to ZnO, implying that the ZnO support is highly crystalline and has a low surface area. Again, no Rh or Rh_2O_3 can be seen on the XRD pattern of Rh/ZnO-800. Fig. S1 in the ESI† depicts the XRD patterns in the $2\theta = 30\text{--}40^\circ$ region, highlighting that there is no difference among the XRD patterns of Rh/Zn- Al_2O_3 -500, Rh/Zn- Al_2O_3 -800, and Rh/ Al_2O_3 -800. No ZnO or ZnAl_2O_4 peaks can be observed for Rh/Zn- Al_2O_3 -500 and Rh/Zn- Al_2O_3 -800, due to the low loading of Zn (*ca.* 1 wt%).

According to the literature,^{34,35} surface spinel (ZnAl_2O_4) may be formed when calcining ZnO/ Al_2O_3 at 800°C . The fact that no ZnAl_2O_4 peaks are observed for Rh/Zn- Al_2O_3 -800 may be because the Zn loading is so low (*ca.* 1 wt%). To prove this explanation, we additionally prepared Zn- Al_2O_3 supports with nominal Zn contents of 0.5 wt%, 5 wt%, 10 wt%, and 20 wt%, respectively. These supports were all calcined at 800°C . All catalysts were accurately scanned in the $2\theta = 30\text{--}40^\circ$ range, where the XRD peaks of γ - Al_2O_3 , ZnAl_2O_4 , and ZnO can be distinguished clearly.

As shown in Fig. 2, the diffraction peaks are significantly enhanced with the increase of Zn content, and gradually shift toward 31.2 and 36.8° which represent (220) and (311) planes of ZnAl_2O_4 , respectively. When the Zn content is or exceeds 10 wt%, the ZnAl_2O_4 peaks become clearer. When the Zn content is 20 wt%, the appearance of peaks at 31.8 , 34.4 , and 36.3° representing (100), (002) and (101) planes of ZnO respectively proves the formation of ZnO. The data infer that surface ZnAl_2O_4 may likely form on Zn- Al_2O_3 -800 with a Zn content of 1 wt%, only that the ZnAl_2O_4 content is very low so the ZnAl_2O_4 phase can not be detected by XRD.

Fig. 3 presents the Raman spectra of several samples without Rh. ZnAl_2O_4 (synthesized *via* citric acid combustion method,³⁷ pure phase proved by XRD) exhibits two Raman peaks at 418 and 659 cm^{-1} , representing E_g and T_{2g} vibration modes of spinel structure, respectively.³⁸ For Zn- Al_2O_3 -800, the peak at

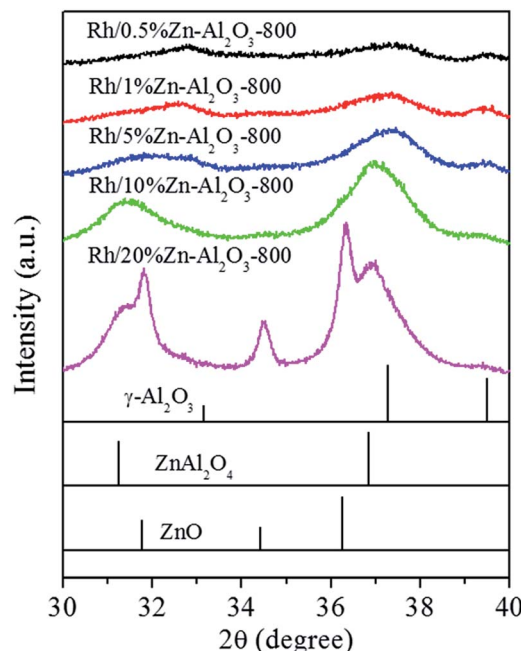


Fig. 2 XRD patterns of Rh/Zn- Al_2O_3 -800 catalysts with different Zn loading and standard patterns of γ - Al_2O_3 (PDF#47-1308), ZnAl_2O_4 (PDF#05-0699), and ZnO (PDF#36-1451).

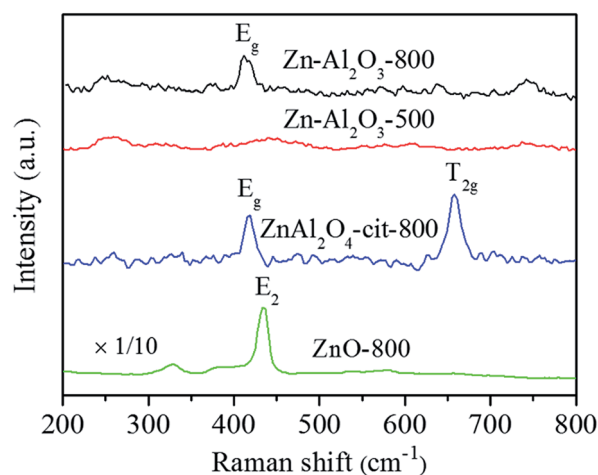


Fig. 3 Raman spectra for Zn- Al_2O_3 -800, Zn- Al_2O_3 -500, and ZnAl_2O_4 -cit-800 (synthesized *via* citric acid route, pure ZnAl_2O_4 phase) and ZnO-800 for comparison.

418 cm^{-1} is obvious, indicating the formation of ZnAl_2O_4 . However, the T_{2g} peak at 659 cm^{-1} is missing, in accordance with a previous report.³⁶ A possible explanation is that tetrahedron defects are induced into spinel structure by impregnating divalent metal cations onto γ - Al_2O_3 (a spinel-like structure with stoichiometric ratio of tetrahedron defects).³⁶ For comparison, Zn- Al_2O_3 -500 does not exhibit peaks at 418 or 659 cm^{-1} . Instead, a broad and weak peak at 440 cm^{-1} is observed, close to E_2 vibration mode of ZnO at 437 cm^{-1} .³⁹ Therefore, high temperature (800°C) is necessary for the formation of surface ZnAl_2O_4 .



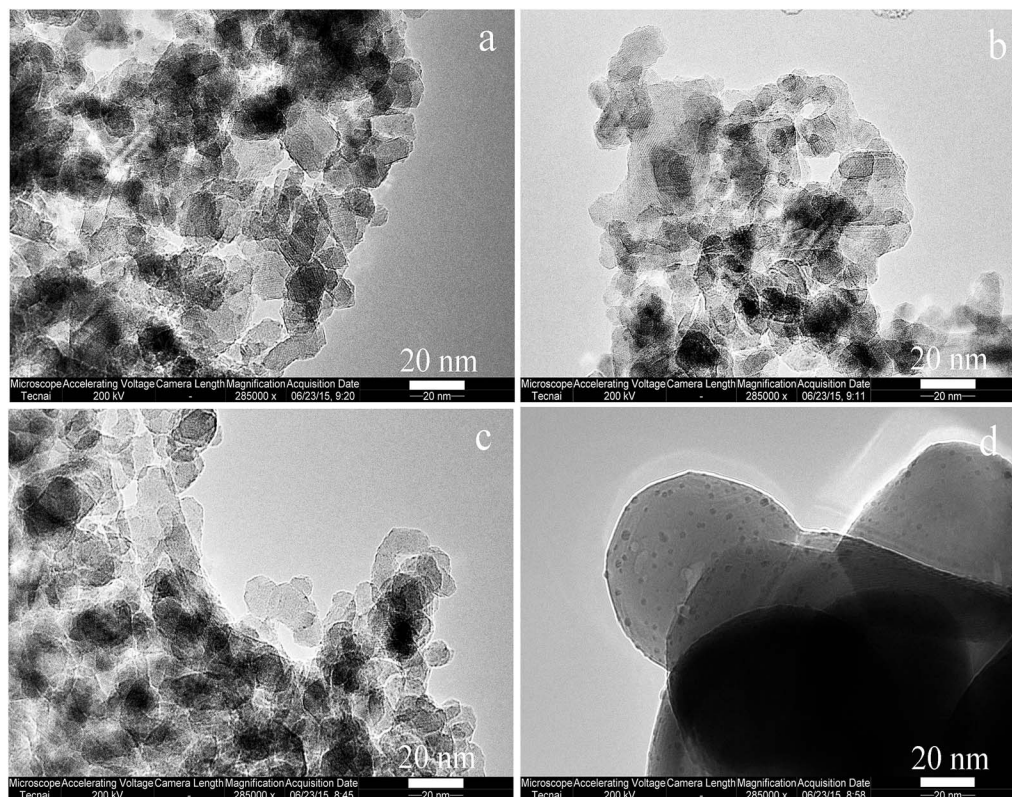


Fig. 4 TEM graphs of (a) Rh/Zn-Al₂O₃-800, (b) Rh/Zn-Al₂O₃-500, (c) Rh/Al₂O₃-800, and (d) Rh/ZnO-800.

Table 1 Some structural and physical properties of Rh/Zn-Al₂O₃ and reference samples

Sample	BET surface area ^a (m ² g ⁻¹)	Average size of Rh ₂ O ₃ particles ^b (nm)	Rh dispersion (%)	Amount of basic sites ^a (μmol g ⁻¹)	Starting desorption temperature of O ₂ (°C)	Amount of desorbed O ₂ (a.u.)
Rh/Zn-Al ₂ O ₃ -800	105	0.70 ± 0.20	71	14.6	234	496
Rh/Zn-Al ₂ O ₃ -500	108	0.78 ± 0.19	65	13.5	293	324
Rh/Al ₂ O ₃ -800	110	0.87 ± 0.25	59	14.3	313	286
Rh/ZnO-800	4	2.15 ± 0.90	7	4.4	—	0

^a Bare supports (without Rh). ^b Obtained by analyzing 300 Rh₂O₃ particles in TEM graphs.

Fig. S2† shows the Raman spectra of Zn-Al₂O₃-800 with different Zn content. The peaks at 418 and 659 cm⁻¹ representing E_g and T_{2g} vibration of ZnAl₂O₄ lattice become stronger as the Zn content in the sample increases from 1 wt% to 5 wt% and then to 10 wt%. The relative intensity of T_{2g} vibration to that of E_g vibration also increases with the Zn content, probably because higher crystallinity of ZnAl₂O₄ restrains the formation of tetrahedron defects. However, when the Zn content of sample is 20 wt%, a new peak at 438 cm⁻¹ representing E₂ vibration of ZnO lattice appears, indicating the formation of ZnO, as also revealed by XRD (Fig. 2).

Fig. 4 shows the TEM graphs of supported Rh catalysts. The size and morphology of Zn-Al₂O₃ supports (Fig. 4a and b) are identical to those of Al₂O₃ (Fig. 4c). This conclusion is in line with the XRD patterns (Fig. 1) and specific surface area data

(Table 1). The surface area of Al₂O₃-800 is 110 m² g⁻¹, and those of Zn-Al₂O₃-500 and Zn-Al₂O₃-800 are 108 and 105 m² g⁻¹, respectively. For comparison, ZnO-800 exhibits as much bigger spherical particles, in accordance with its low surface area (4 m² g⁻¹, Table 1). Rh₂O₃ particles are dispersed on these supports. The homogeneous distribution of Rh and Zn on Rh/Zn-Al₂O₃-800 is shown by EDX-mapping (Fig. S3†).

Fig. S4† shows the TEM graphs with larger size. The graphs show that Rh₂O₃ are well dispersed on the supports. The size distributions of Rh₂O₃ particles were determined by analyzing 300 particles from over 5 TEM graphs for each sample. As shown in Fig. S5,† most Rh₂O₃ particles are in the range of 0.5 and 1.0 nm over Al₂O₃-based catalysts. However, due to the low surface area of ZnO support (4 m² g⁻¹), Rh₂O₃ particles on ZnO are relatively big. The average sizes of Rh₂O₃ particles on Zn-



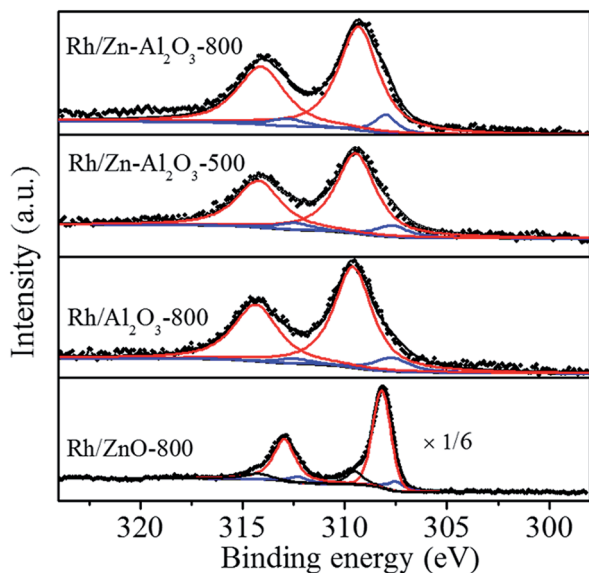


Fig. 5 Rh 3d XPS spectra of Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800, and Rh/ZnO-800.

Al₂O₃-800, Zn-Al₂O₃-500, Al₂O₃-800, and ZnO-800 are 0.70, 0.78, 0.87, and 2.15 nm, respectively. Rh dispersions of Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800, and Rh/ZnO-800 are 71%, 65%, 59%, and 7%, respectively, in line with the trend of Rh₂O₃ particle size seen by TEM. The key observation here is that the addition of ZnO onto commercial Al₂O₃ can stabilize Rh₂O₃ particles, and the calcination of ZnO/Al₂O₃ at 800 °C can exert such an effect more obviously.

Fig. 5 shows the Rh 3d XPS spectra of fresh catalysts. Peaks assigned to Rh 3d_{5/2} and Rh 3d_{3/2} are at 306–311 eV and 311–317 eV, respectively. The following discussion will be focused on the Rh 3d_{5/2} peak due to its higher intensity. The binding energy of reduced Rh species (Rh⁰) is at 307.0–307.7 eV, that of non-stoichiometric Rh oxide (Rh⁺) is at about 308.1 eV, and that of Rh³⁺ is at 308.3–310.5 eV.^{6,12,40–42} Rh 3d_{5/2} peaks are located at 309.3, 309.5, and 309.6 eV for Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, and Rh/Al₂O₃-800, respectively, indicating that Rh species mainly exist as Rh³⁺ in the form of Rh₂O₃. Only a small portion of Rh species (<10%) exist as Rh⁰. It is reasonable because the catalysts have been calcined at 500 or 800 °C in air and without further reduction. However, Rh 3d spectrum of Rh/ZnO-800 is quite different, showing a very sharp and strong peak at 308.2 eV, indicating the formation of non-stoichiometric Rh oxide. Due to low surface area of ZnO, relatively big Rh₂O₃ particles on ZnO (Fig. S4d†) have different chemical properties as compared to smaller particles.¹¹ In addition, high surface Rh density on low-surface-area ZnO results in stronger Rh signal. Fig. S6† compares the XPS spectra of Rh/Zn-Al₂O₃-800 with and without being pretreated in 4% H₂ (balance He) at 400 °C for 2 h. Binding energy of Rh 3d_{5/2} peak declines from 309.3 eV (without pretreatment) to 308.6 eV (with pretreatment), indicating the partial reduction of Rh species (the proportion of Rh⁰ increases from 9.1% to 26.0%) upon H₂ pretreatment.

Activity measurement

Fig. 6 shows the N₂O conversions on supported Rh catalysts as a function of reaction temperature. The catalytic activity on these catalysts follows the sequence of Rh/Zn-Al₂O₃-800 > Rh/Zn-Al₂O₃-500 > Rh/Al₂O₃-800 > Rh/ZnO-800. The N₂O conversions over these catalysts at 275 °C are 98.7%, 54.4%, 27.2%, and 22.3%, respectively. The specific rates of these catalysts (expressed as moles of N₂O converted per mole of Rh per minute) at 275 °C are calculated to be 0.226, 0.125, 0.069, and 0.056 min⁻¹, respectively. The T₅₀ (temperature required for 50% conversion) values of these catalysts are 251, 273, 289, and 300 °C, respectively. Note that ZnO referred to above was prepared by precipitation. An additional Rh/ZnO-800 catalyst was prepared using ZnO obtained by calcining Zn(NO₃)₂·6H₂O at 800 °C for 4 h. That catalyst is less active than Rh/ZnO-800 mentioned above (Fig. S7†). In addition, Zn-Al₂O₃-800 (without Rh) is much less active (17.0% N₂O conversion at 400 °C) than Rh/Zn-Al₂O₃, indicating that Rh₂O₃ act as the main active sites for N₂O decomposition.

The effect of calcination temperature of Zn-Al₂O₃ on the activity of the resulting Rh/Zn-Al₂O₃ catalysts was studied. As shown in Fig. S8,† the N₂O conversions over Rh/Zn-Al₂O₃-500, Rh/Zn-Al₂O₃-600, Rh/Zn-Al₂O₃-700, Rh/Zn-Al₂O₃-800, and Rh/Zn-Al₂O₃-900 at 275 °C are 54.4%, 55.5%, 68.0%, 98.7%, and 47.1%, respectively. The T₅₀ values of these catalysts are 273, 272, 266, 251, and 277 °C, respectively. The activities of these catalysts follow the sequence of Rh/Zn-Al₂O₃-900 < Rh/Zn-Al₂O₃-500 ~ Rh/Zn-Al₂O₃-600 < Rh/Zn-Al₂O₃-700 < Rh/Zn-Al₂O₃-800, i.e., Rh/Zn-Al₂O₃-800 is the most active.

The effect of Zn content on the performance of Rh/Zn-Al₂O₃ catalysts was also studied. As shown in Fig. S9,† the N₂O conversions over Rh/Al₂O₃-800, Rh/0.5% Zn-Al₂O₃-800, Rh/1% Zn-Al₂O₃-800, Rh/10% Zn-Al₂O₃-800, and Rh/20% Zn-Al₂O₃-800 at 275 °C are 27.2%, 89.7%, 98.7%, 83.7%, and 70.1%, respectively. The T₅₀ values of these catalysts are 289, 258, 251, 261, and 265 °C, respectively. The activities of these catalysts follow the sequence of Rh/Al₂O₃-800 < Rh/20% Zn-Al₂O₃-800 < Rh/10% Zn-Al₂O₃-800 < Rh/0.5% Zn-Al₂O₃-800 < Rh/1% Zn-Al₂O₃-800, i.e., Rh/1% Zn-Al₂O₃-800 is the most active.

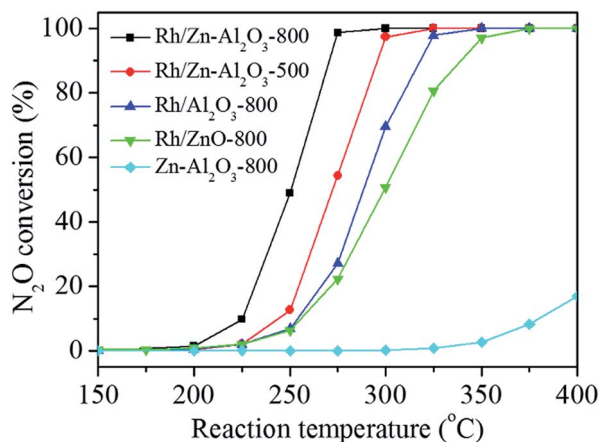


Fig. 6 Conversion of N₂O on the catalysts as a function of reaction temperature.

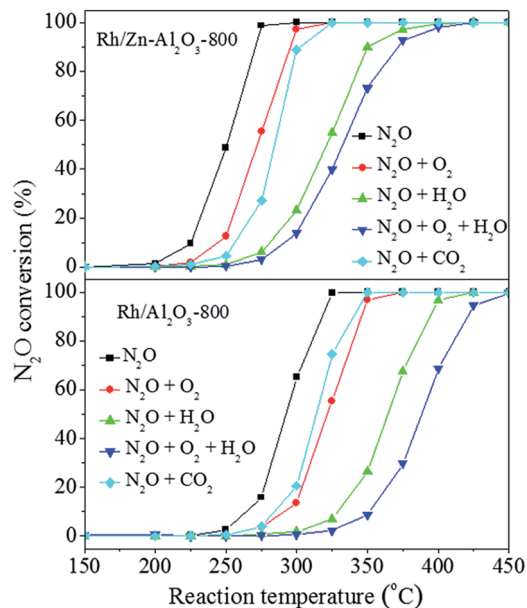


Fig. 7 Influence of co-feed 5% O_2 , 2% H_2O , or 2% CO_2 on the conversion of N_2O over $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$ and $\text{Rh}/\text{Al}_2\text{O}_3-800$ as a function of reaction temperature.

Fig. S10† shows the influence of different GHSV values (7420, 14 840, and 37 100 h^{-1}), achieved by adjusting the catalyst dosage (0.5, 0.25, and 0.1 g, respectively), on the performance of the optimal catalyst $\text{Rh}/1\% \text{Zn}-\text{Al}_2\text{O}_3-800$. In general, a decrease of N_2O conversion is seen at a high GHSV at the same reaction temperature. The T_{50} values are 251, 266, and 291 $^{\circ}\text{C}$ at GHSV of 7420, 14 840, and 37 100 h^{-1} , respectively. However, the N_2O conversion still reaches 87.5% at 325 $^{\circ}\text{C}$ when the GHSV is as high as 37 100 h^{-1} .

Fig. S11† shows the influence of $\text{Rh}/\text{M}-\text{Al}_2\text{O}_3-800$ ($\text{M} = \text{Zn}, \text{Mg}, \text{Co}, \text{Ni}, \text{Cu}$) catalysts with different divalent metal cations on N_2O decomposition. The M content is fixed to be 1 wt%. All $\text{Rh}/\text{M}-\text{Al}_2\text{O}_3-800$ catalysts exhibit superior performance than $\text{Rh}/\text{Al}_2\text{O}_3-800$. The N_2O conversions on $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$, $\text{Rh}/\text{Mg}-\text{Al}_2\text{O}_3-800$, $\text{Rh}/\text{Co}-\text{Al}_2\text{O}_3-800$, $\text{Rh}/\text{Ni}-\text{Al}_2\text{O}_3-800$, $\text{Rh}/\text{Cu}-\text{Al}_2\text{O}_3-800$, and $\text{Rh}/\text{Al}_2\text{O}_3-800$ at 275 $^{\circ}\text{C}$ are 98.7%, 64.1%, 76.0%, 72.9%, 82.5%, and 27.2% respectively. $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$ is still the most active among these catalysts.

Fig. 7 shows N_2O conversion over $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$ and $\text{Rh}/\text{Al}_2\text{O}_3-800$ in the absence or presence of H_2O , O_2 , or CO_2 . According to the literature,^{6,19} H_2O and O_2 cause competitive adsorption with N_2O on active sites, leading to severe inhibition of activity on Rh -based catalysts. The T_{50} value of $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$ in the absence of H_2O , O_2 , or CO_2 is 251 $^{\circ}\text{C}$. The T_{50} value increases to 272 $^{\circ}\text{C}$ when 5% O_2 alone is added to the reaction mixture, and it increases more obviously to 321 $^{\circ}\text{C}$ when 2% H_2O alone is added. When 5% O_2 and 2% H_2O are co-fed to the reaction mixture, the T_{50} value increases further to 333 $^{\circ}\text{C}$. The addition of 2% CO_2 alone makes the T_{50} value become 284 $^{\circ}\text{C}$. The presence of H_2O , O_2 , or CO_2 also exerts similar inhibiting effects on $\text{Rh}/\text{Al}_2\text{O}_3-800$, only that $\text{Rh}/\text{Al}_2\text{O}_3-800$ is always much less active than $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$ under respective condition for

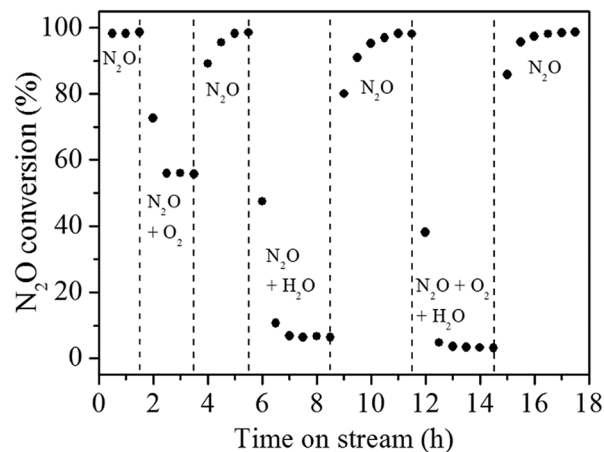


Fig. 8 Response of N_2O conversion to step changes of co-feed 5% O_2 and/or 2% H_2O over $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$ at 275 $^{\circ}\text{C}$.

comparison. The stability of both $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$ and $\text{Rh}/\text{Al}_2\text{O}_3-800$ as a function of time on stream is good, either in the absence or presence of 2% H_2O and 5% O_2 (Fig. S12†).

Gas-switching experiments were conducted to know whether the inhibiting effect of O_2 , and H_2O is reversible. Fig. 8 clearly shows that N_2O conversion decreases almost immediately when 5% O_2 is co-fed into the reactor, but the conversion can be restored when stopping feeding O_2 . 2% H_2O has a more obviously inhibiting effect, but such an inhibiting effect is also reversible.^{11,26}

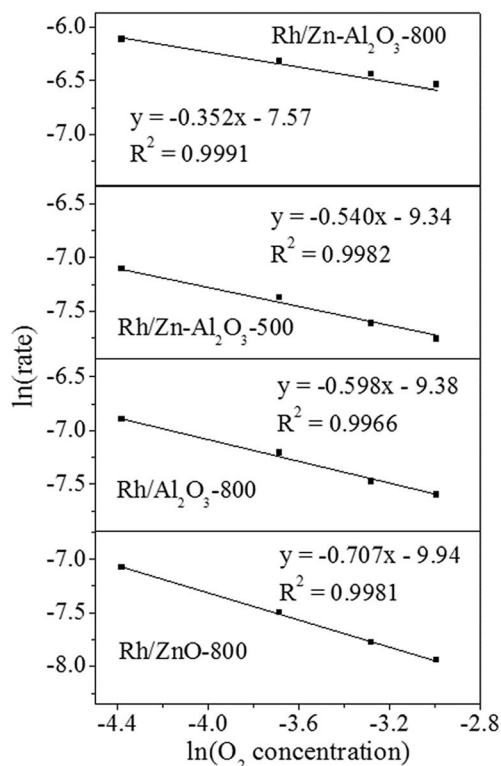


Fig. 9 Influence of O_2 concentration on the activities of $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-800$, $\text{Rh}/\text{Zn}-\text{Al}_2\text{O}_3-500$, $\text{Rh}/\text{Al}_2\text{O}_3-800$, and $\text{Rh}/\text{ZnO}-800$ in N_2O decomposition.



O₂ reaction order was measured by changing O₂ concentration while keeping N₂O concentration as 0.5%,⁴³ to further explore the inhibition effect of oxygen. To eliminate diffusion factor, catalysts usage (0.128, 0.204, 0.256, and 0.306 g for Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800 and Rh/ZnO-800, respectively) and reaction temperature (230, 250, 260, and 290 °C for Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800 and Rh/ZnO-800, respectively) were chosen to control the N₂O conversion below 15%. As shown in Fig. 9, a good linearity between ln(rate) versus ln(O₂ concentration) enables us to derive O₂ reaction orders as −0.352, −0.540, −0.598, and −0.707 for Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800 and Rh/ZnO-800, respectively. The data indicate that the inhibition effect of O₂ on Rh/Zn-Al₂O₃-800 is the weakest, whereas that on Rh/ZnO-800 is the strongest.

The influence of 4% H₂ pretreatment on Rh/Zn-Al₂O₃-800 was investigated. As shown in Fig. S13,[†] the H₂ pretreatment could improve the activity. The N₂O conversion reaches 64.5% at 250 °C (in comparison to the 48.9% conversion achieved on the catalyst without being pretreated in H₂). It has been reported that metallic Rh is more favorable for N₂O decomposition on non-reducible support,^{10,25} thus the formation of some metallic Rh after H₂ reduction (the proportion of Rh⁰ increases from 9.1% to 26.0% upon H₂ pretreatment, Fig. S6[†]) should be the reason of the promoted activity although Rh₂O₃ can also catalyze this reaction.

Chemical properties of Rh/Zn-Al₂O₃ and reference samples

Fig. S14[†] shows the CO₂-TPD profiles of Zn-Al₂O₃-800, Zn-Al₂O₃-500, Al₂O₃-800, and ZnO-800, and the values of calculated basicity are listed in Table 1. It was reported that ZnAl₂O₄ exhibits higher basic site density than Al₂O₃.⁴⁴ However, here Zn-Al₂O₃-800, Zn-Al₂O₃-500, and Al₂O₃-800 show similar CO₂-TPD profiles. The amounts of basic sites of these supports are 14.6, 13.5, and 14.3 μmol g^{−1}, respectively. Low Zn loading should be the reason for the similarity, because only a small amount of ZnAl₂O₄ phase is formed on Al₂O₃ surface. On the

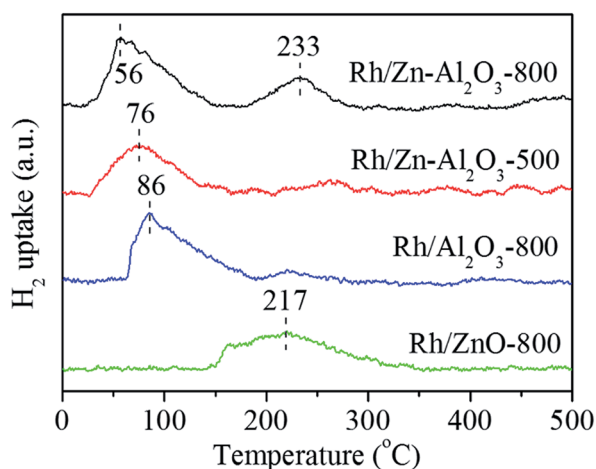


Fig. 10 H₂-TPR profiles of Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800 and Rh/ZnO-800 catalysts.

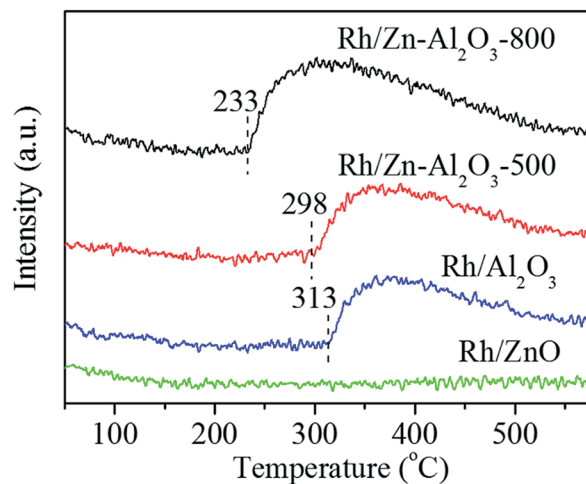


Fig. 11 O₂-TPD profiles of Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800 and Rh/ZnO-800 catalysts.

other hand, ZnO-800 with a low surface area (4 m² g^{−1}) has the smallest amount of basic sites (4.4 μmol g^{−1}), in accordance with a previous report.⁴⁵

Fig. 10 shows the H₂-TPR profiles of Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800, and Rh/ZnO-800. All catalysts exhibit reduction peaks below 300 °C, assigned to the reduction of Rh species.^{17,30,42} The reduction temperature follows the sequence of Rh/Zn-Al₂O₃-800 (56 °C) < Rh/Zn-Al₂O₃-500 (76 °C) < Rh/Al₂O₃ (86 °C) < Rh/ZnO-800 (217 °C). The reduction temperature of Rh/ZnO-800 is much higher than those of other catalysts, probably due to larger Rh particles (TEM, Fig. 4 and S4[†]) with a distinctive electronic property (XPS, Fig. 5). In addition, a peak located at 233 °C is exhibited (the result is reproducible) in Rh/Zn-Al₂O₃-800, may due to the strong interaction between the support and Rh, resulting in the formation of other Rh-containing species instead of Rh₂O₃.^{46,47}

Fig. 11 shows the O₂-TPD profiles of Rh/Zn-Al₂O₃-800, Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800, and Rh/ZnO-800. Rh/ZnO-800 does not show any oxygen desorption peak, whereas other three catalysts all show an oxygen desorption peak below 580 °C. The starting desorption temperature follows the order of Rh/Zn-Al₂O₃-800 (233 °C) < Rh/Zn-Al₂O₃-500 (298 °C) < Rh/Al₂O₃-800 (313 °C), whereas the amounts of oxygen desorption follow the sequence of Rh/Zn-Al₂O₃-800 > Rh/Zn-Al₂O₃-500 > Rh/Al₂O₃-800 (Table 1).

Discussion

The N₂O decomposition activities of four catalysts follow the sequence of Rh/Zn-Al₂O₃-800 > Rh/Zn-Al₂O₃-500 > Rh/Al₂O₃-800 > Rh/ZnO-800 (Fig. 6). The catalytic performance of the best catalyst (Rh/Zn-Al₂O₃-800) in this work is then compared with the performance of other catalysts. Table S1[†] lists the *T*₅₀ and *T*₉₀ values of various catalysts, together with the reaction conditions. Although the Rh/Zn-Al₂O₃-800 catalyst seems to be more active than most of the catalysts, such as Cu-,^{48–50} Fe-,^{51,52} and Co-based catalysts,^{53–56} as well as Pd-, Pt-,^{8,10,57} and Ir-based⁵⁸



catalysts (judging from the T_{50} values), it is not appropriate to claim that because of the difference in reaction conditions. Hence, the activity of Rh/Zn-Al₂O₃-800 is compared with Rh/SiO₂, Rh/Al₂O₃, Rh/TiO₂, and Rh/CeO₂ (all with *ca.* 1 wt% Rh) tested under the same reaction condition in our laboratory.²⁵ As shown in Table S1,† the T_{50} values of Rh/Zn-Al₂O₃-800, Rh/SiO₂, Rh/Al₂O₃, Rh/TiO₂, and Rh/CeO₂ are 251, 324, 289, 310, and 223 °C, respectively, indicating the good performance of Rh/Zn-Al₂O₃-800 among these supported Rh catalysts. The activity sequence of catalysts is Rh/CeO₂ > Rh/Zn-Al₂O₃-800 > Rh/Al₂O₃ > Rh/TiO₂ > Rh/SiO₂. Rh/Zn-Al₂O₃-800 (T_{50} = 251 °C) is less active than Rh/HAP-PEG-200 (*ca.* 1 wt% Rh supported on hydroxyapatite nanorods synthesized through hydrothermal method, and tested under the same reaction condition) that shows a T_{50} value of 223 °C.²⁷ However, the synthesis of Rh/Zn-Al₂O₃-800 is more convenient and of low cost, and Rh/Zn-Al₂O₃-800 shows superior activity than Rh/HAP-PEG-200 in the presence of 2% H₂O and 5% O₂. T_{50} values of these two catalysts under such a condition are 333 and 344 °C, respectively.

After comparing the catalytic activity, we then attempt to discuss why Rh/Zn-Al₂O₃-800 is more active than Rh/Zn-Al₂O₃-500, Rh/Al₂O₃-800, and Rh/ZnO-800. The structural properties of γ -Al₂O₃ are not significantly altered by the incorporation of 1 wt% Zn, as seen from the micro-morphology (TEM, Fig. 4) and specific surface areas (Table 1). The average size of Rh₂O₃ particles follows the sequence of Rh/Zn-Al₂O₃-800 < Rh/Zn-Al₂O₃-500 < Rh/Al₂O₃-800 < Rh/ZnO-800 (also proved by Rh dispersion, Table 1), inversely correlating to the order of catalytic activity, in the sense that a catalyst with smaller Rh₂O₃ particles has higher activity in N₂O decomposition.^{25,26,28} The data thus infer that the formation of surface ZnAl₂O₄ spinels on Al₂O₃ support may be beneficial for the high dispersion of Rh₂O₃ particles.

In our ongoing research, bulk ZnAl₂O₄ support is also found to be able to lead to high dispersion of Rh₂O₃ species, thus promoting the catalytic activity (data not shown). There seems to be a stronger interaction between Rh₂O₃ and ZnAl₂O₄ phase. Similar conclusions can be found in the literature. As revealed by theoretical calculation reveals that MgAl₂O₄ (spinel structure) has a stronger interaction with supported Rh and Ir particles than γ -Al₂O₃, promoting the dispersion of Rh and Ir as well as the stability of the catalysts in methane steam reforming.³² In addition, smaller Pd nanoparticles could be immobilized on ZnAl₂O₄ surface, leading to higher activity and good stability in Suzuki-Miyaura coupling reaction.³³ It was also reported that Rh prefers to be located on the spinel phase of metal modified-Al₂O₃ support.⁵⁹ Therefore, the better Rh dispersion on Rh/Zn-Al₂O₃-800 should be caused by the formation of ZnAl₂O₄ phase on Al₂O₃ support, as proved by Raman (Fig. 3). XRD data also provide indirect evidence because Zn-Al₂O₃-800 supports with low Zn contents do not show ZnAl₂O₄ peaks due to the low concentration of Zn, whereas Zn-Al₂O₃-800 supports with relatively high Zn contents show ZnAl₂O₄ peaks clearly (Fig. 2). Another piece of evidence showing the strong interaction between Rh₂O₃ and ZnAl₂O₄ comes from the H₂-TPR data, as an additional reduction peak shows up at a higher temperature of 233 °C (Fig. 10).

As shown in XPS data (Fig. 5), the binding energies of Rh 3d_{5/2} follow the sequence of Rh/Zn-Al₂O₃-800 (309.3 eV) < Rh/Zn-Al₂O₃-500 (309.5 eV) < Rh/Al₂O₃-800 (309.6 eV), indicating the electronic properties of Rh species are altered by the supports. Rh/Zn-Al₂O₃-800 exhibits the lowest Rh 3d_{5/2} binding energy among these catalysts, indicating the strongest shielding effect caused by the highest outer electron density. It has been reported that metal-oxygen bonding energy is weaker in metal oxides with higher electron density, because the electron could fill into the empty orbit of oxygen.⁶⁰⁻⁶³ As a result, Rh₂O₃ on Zn-Al₂O₃-800 is easier to be reduced due to the weaker bonding between Rh and oxygen (H₂-TPR, Fig. 10). In the literature, the higher activity of catalysts in N₂O decomposition has been correlated to the reduction behaviour of catalysts (more specifically, the active components) in some cases.⁶⁴⁻⁶⁸ That is not to say that these catalysts have to be reduced prior to the reaction testing. Rather, the enhanced reducibility means the weakening of the M-O bonds, which is expected to be important for N₂O decomposition, *i.e.*, the adsorbed oxygen can be desorbed from the catalyst more easily.

In another aspect, oxygen desorption is a key step in N₂O decomposition, even rate-determining step in many cases.^{2,10,24,56,69} If oxygen can not be smoothly desorbed from catalysts surface, the active sites will be occupied by oxygen, blocking the catalytic circle. As shown in Fig. 9, the starting O₂-desorption temperature follows the order of Rh/Zn-Al₂O₃-800 (233 °C) < Rh/Zn-Al₂O₃-500 (298 °C) < Rh/Al₂O₃-800 (313 °C), in line with the notion that easy desorption of O₂ is beneficial for catalytic N₂O decomposition.^{26,27,70,71} In addition, the inhibition effect of O₂ (as seen from the oxygen reaction order, Fig. 8) on catalytic activity follows the sequence of Rh/Zn-Al₂O₃-800 < Rh/Zn-Al₂O₃-500 < Rh/Al₂O₃-800 < Rh/ZnO, inversely correlating with the activity sequence of these catalysts.

Conclusions

Zn(NO₃)₂ was impregnated onto a commercial γ -Al₂O₃ support. The composite was then calcined in air at 500 or 800 °C. The resulting Zn-Al₂O₃-500 and Zn-Al₂O₃-800 supports were used to load Rh *via* impregnation with Rh(NO₃)₃ and calcination at 500 °C. While the Zn species in Rh/Zn-Al₂O₃-500 is ZnO, the Zn species in Rh/Zn-Al₂O₃-800 is ZnAl₂O₄. The activities of four catalysts in N₂O decomposition follow the sequence of Rh/Zn-Al₂O₃-800 > Rh/Zn-Al₂O₃-500 > Rh/Al₂O₃-800 > Rh/ZnO-800, correlating with the size of Rh₂O₃ particles, the reducibility of Rh₂O₃, the O₂-desorption property, and the oxygen reaction order. The most active Rh/Zn-Al₂O₃-800 has the smallest Rh₂O₃ particles (*i.e.*, the highest Rh dispersion), the lowest reduction temperature of Rh₂O₃, the lowest O₂ desorption temperature, and its activity is influenced by the presence of O₂ least obviously. The formation of ZnAl₂O₄ on Al₂O₃ support is beneficial for the stabilization of Rh₂O₃ particles. This work demonstrates a convenient way for preparing active Rh catalysts based on commercial γ -Al₂O₃ support.

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