PAPER

Aggregation induced emission and amplified explosive detection of tetraphenylethylene-substituted polycarbazoles

Wenyue Dong,[a] Teng Fei,[b] Alex Palma-Cando[a] and Ullrich Scherf*[a]

Novel conjugated polymers based on 3,6-carbazole repeat units were synthesized by nickel-catalyzed Yamamoto coupling under microwave heating. The resulting poly(3,6-carbazole)s contain tetraphenylethylene (TPE) units in their side chains. The resultant polymers show aggregation induced emission (AIE) behavior. Hereby, the photoluminescence (PL) intensity of PCzTPE0.5 in 90% water–THF is 35 times higher than that in pure THF, connected to the introduction of TPE side chains. The ability of polymer PCzTPE0.5 for explosive sensing was also studied. A maximum Stern–Volmer quenching constant of 1.26×10^6 M^{-1} was observed for PL quenching of PCzTPE0.5 aggregates by trinitrobenzene (TNB). A solid state paper strip test based on PCzTPE0.5 and PCzTPE also demonstrates effective PL quenching towards both TNB vapor and solution.

Introduction

Most "conventional" organic luminophores exhibit high photoluminescence (PL) efficiency in solution but weaker PL efficiency in the solid state (*e.g.* in the films). The effect is called aggregation caused quenching (ACQ)[1–3] and is often detrimental for practical applications, *e.g.* in sensor devices. To solve this problem, several chemical, physical or engineering approaches[4–11] have been developed, but with limited success, because one has to fight against an intrinsic process – the energetically favored formation of chromophore aggregates in the solid state. In 2001, Tang *et al.* first described a novel phenomenon opposite to the ACQ effect: propeller-shaped molecules such as hexaphenylsilole (HPS) and tetraphenylethylene (TPE) emit strongly in the solid state but are weakly emissive in solution. The phenomenon is called aggregation induced emission (AIE).[12–14] Thanks to the AIE effect, the design of efficient solid emitters was possible. The main reason for the AIE behavior is the restriction of intramolecular rotation (RIR) of the peripheral phenyl rotors in the aggregated state as has been concluded both from experimental and theoretical studies.[15–19]

The detection of explosives, such as nitroaromatic compounds or peroxides, becomes increasingly important in modern society, due to the concerns on global security and environmental protection.[20–23] Many methods have been

developed for explosive detection, such as gas chromatography,[24] mass spectrometry,[25,26] surface enhanced Raman spectroscopy,[27] ion mobility spectrometry,[28,29] electrochemical sensing,[30] PL spectroscopy,[31–34] and others. Among them, PL sensors based on conjugated polymers[35] have been widely tested because of their simplicity and high sensitivity. Trinitroaromatic compounds containing three electron-withdrawing nitro groups are potent electron acceptors. The working principle of PL sensors for nitroaromatic compounds is based on a photo-induced electron transfer from donor (conjugated polymer) to acceptor (nitroaromatic compound) thus resulting in PL quenching. Recently, AIE luminogens have been used in PL sensors for the detection of nitroaromatics[36–42] due to their high solid state fluorescence quantum yields. Moreover, their twisted structure creates an increased number of 3D exciton diffusion channels, thus enhancing the quenching efficiency.

In this study, we prepared novel AIE-active conjugated polymers and investigated their ability for sensing nitroaromatic compounds. Two carbazole-based polymers (PCzTPE and PCzTPE0.5) with TPE side chain units were successfully synthesized. The 3,6-carbazole repeat units[43–45] were chosen for constructing the polycarbazole backbone since their electron-donor character should be beneficial for the interaction with electron-poor trinitroaromatics. The incorporation of TPE into the side chains guarantees AIE activity, without strongly affecting the electronic properties of the polycarbazole backbone. Based on this design principle we expected a high sensitivity for the detection of nitroaromatic compounds. Our novel polymers are the first polymeric AIE materials with the AIE active groups in the side chain.

[a]Bergische Universität Wuppertal, Macromolecular Chemistry Group (buwmakro) and Institute for Polymer Technology, Gauss-Str. 20, D-42097 Wuppertal, Germany. E-mail: Scherf@uni-wuppertal.de

[b]State Key Laboratory on Integrated Optoelectronics, College of Electronic Science and Engineering, Jilin University, Changchun 130012, PR China

Experimental

Characterization of materials

NMR spectra were recorded on a Bruker AVANCE 400 or AVANCE III 600. ^{1}H and ^{13}C NMR spectra were recorded with tetramethylsilane (TMS) as an internal standard. Gel permeation chromatography (GPC) measurements were carried out on a PSS/Agilent SECurity GPC System equipped with polystyrene gel columns using chloroform as an eluent. APLI (Atmospheric Pressure Laser Ionization) measurements were carried out on a Bruker Daltronik Bremen with micrOTOF. UV-visible absorption spectra were recorded on a Jasco V-670 spectrometer and PL spectra on a Varian CARY Eclipse F2500. Elemental analyses were performed on a Vario EL II (CHNS) instrument. Thermal gravimetric analysis (TGA) was undertaken on a TGA/DSC1 STAR System (Mettler Toledo) at a heating rate of 10 °C min^{-1} and an argon flow rate of 50 mL min^{-1}. Differential scanning calorimetry (DSC) was performed on a DSC1 STAR System (Mettler Toledo) at a heating rate of 10 °C min^{-1} under argon. The PL quantum efficiencies of polymer films were measured with an integrating sphere. Cyclic voltammetry (CV) measurements of the polymer films were performed on a standard three-electrode electrochemical cell attached to a VersaSTAT 4 electrochemical workstation in dichloromethane for polymers and acetonitrile for trinitrobenzene (TNB) with 0.1 M tetrabutylammonium perchlorate as a supporting electrolyte at a scan rate of 0.1 V s^{-1} for polymers and 0.2 V s^{-1} for TNB. The potentials were measured against an Ag/AgNO$_3$ reference electrode (0.1 M AgNO$_3$ in acetonitrile/0.6 V vs. NHE). The onset potentials were determined from the intersection of two tangents drawn at the rising current and background current of the cyclic voltammogram.

Synthesis

All reagents were obtained from commercial suppliers and were used without further purification. All reactions were carried out under an argon atmosphere using standard and Schlenk techniques. The solvents used were of commercial p.a. quality.

1-(4-Fluorophenyl)-1,2,2-triphenylethylene (1). To a solution of diphenylmethane (8.08 g, 48 mmol) in dry tetrahydrofuran (80 mL) 2.8 M solution of *n*-butyllithium in hexane (48 mmol) was added at 0 °C under an argon atmosphere. The resulting orange-red solution was stirred for 1 h at that temperature. To this solution 4-fluorobenzophenone (8.00 g, 40 mmol) was added, and the reaction mixture was allowed to warm up to room temperature overnight. The reaction was quenched with the addition of an aqueous solution of ammonium chloride. The organic layer was extracted with chloroform, and the combined organic layers were washed with a saturated brine solution and dried over anhydrous MgSO$_4$. The solvent was evaporated, and the resulting crude alcohol (containing excess diphenylmethane) was subjected to acid-catalyzed dehydration as follows.

The crude alcohol was dissolved in about 250 mL of toluene containing *p*-toluenesulphonic acid (2.0 g, 10.5 mmol) in a 500 mL flask, and the mixture was refluxed overnight. The toluene layer was washed with 10% aqueous NaHCO$_3$ solution and dried over MgSO$_4$ and evaporated to afford the crude tetraphenylethylene

derivative. The crude product was purified by recrystallization from the mixture of dichloromethane and methanol to give the target compound as a white solid in 71% yield (10.0 g). ^{1}H NMR (400 MHz, C$_2$D$_2$Cl$_4$) δ 7.15–7.07 (m, 9H), 7.06–6.95 (m, 8H), 6.80 (t, J = 8.8 Hz, 2H). ^{13}C NMR (C$_2$D$_2$Cl$_4$, 100 MHz) δ (ppm) 160.24, 143.73, 143.69, 143.62, 141.41, 140.01, 139.97, 139.94, 133.23, 133.15, 131.51, 128.06, 128.00, 127.95, 126.76, 126.69, 114.94, 114.73.

1-[4-(3,6-Dibromocarbazole-9-yl)phenyl]-1,2,2-triphenylethylene (2). A solution of compound **1** (3.00 g, 8.56 mmol), 3,6-dibromocarbazole (3.06 g, 9.42 mmol) and K$_3$PO$_4$ (9.08 g, 42.8 mmol) in DMF (120 mL) was stirred at 150 °C for 24 h under an argon atmosphere. The reaction mixture was quenched with water and extracted with chloroform. The organic phases were collected, dried over MgSO$_4$, and concentrated in a vacuum. The product was purified by silica gel chromatography (eluent: hexane–dichloromethane = 4/1) to give the desired compound as a white solid in 25% yield (1.4 g). ^{1}H NMR (400 MHz, C$_2$D$_2$Cl$_4$) δ 8.19 (d, J = 1.9 Hz, 2H), 7.53 (dd, J = 8.7, 1.9 Hz, 2H), 7.26 (d, J = 8.6 Hz, 2H), 7.24–7.12 (m, 15H), 7.12–7.06 (m, 4H). ^{13}C NMR (C$_2$D$_2$Cl$_4$, 100 MHz) δ 143.94, 143.66, 143.42, 143.20, 142.42, 140.13, 139.99, 134.69, 133.20, 131.66, 131.60, 131.52, 129.69, 128.19, 128.05, 128.03, 127.04, 126.99, 126.10, 124.06, 123.44, 113.26, 111.94. MS (APLI): m/z calcd 655.03; found 655.03. Elemental anal. calcd for **2**: C, 69.64%; H, 3.84%; N, 2.14%. Found: C, 69.17%; H, 3.84%; N, 2.16%.

3,6-Dibromo-9-octylcarbazole (3). 3,6-Dibromocarbazole (5.0 g, 1.55 mmol) and tetrabutylammonium bromide (TBABr) (500 mg, 1.55 mmol) were dissolved in 100 mL of DMSO under argon. Aqueous NaOH solution (1 g mL^{-1}, 8 mL) was added to the mixture, and the mixture was stirred at 60 °C for 5 min. Then, 1-bromooctane (4 mL, 0.023 mol) and 10 mL of DMSO were added and the reaction mixture was heated to 90 °C overnight. The reaction mixture was quenched with water and extracted with chloroform. The organic phases were collected, dried over MgSO$_4$ and concentrated under vacuum. The product was purified by silica gel chromatography (eluent: hexane–dichloromethane = 3/1) to give the desired compound as a white solid in 81% yield (5.4 g). ^{1}H NMR (400 MHz, C$_2$D$_2$Cl$_4$) δ 8.15 (s, 2H), 7.58 (dd, J = 8.7, 1.9 Hz, 2H), 7.30 (d, J = 8.7 Hz, 2H), 4.23 (t, J = 7.0 Hz, 2H), 1.88–1.74 (m, 2H), 1.38–1.15 (m, 10H), 0.87 (t, J = 6.8 Hz, 3H). ^{13}C NMR (C$_2$D$_2$Cl$_4$, 100 MHz) δ (ppm) 139.57, 129.37, 123.58, 123.49, 112.16, 110.88, 43.69, 32.03, 29.58, 29.40, 29.15, 27.49, 22.91, 14.48. MS (APLI): m/z calcd 437.02; found 437.06. Elemental anal. calcd for **3**: C, 54.94%; H, 5.30%; N, 3.20%. Found: C, 54.91%; H, 5.58%; N, 3.20%.

Polymer PCzTPE. A solution of compound **2** (400 mg, 0.610 mmol), Ni(COD)$_2$ (436 mg, 1.587 mmol), BPy (110 mg, 0.701 mmol) and COD (172 mg, 1.587 mmol) in 7 mL of THF was reacted under microwave heating at 120 °C for 12 min. The reaction mixture was quenched with water and extracted with chloroform. The collected organic phases were washed with aqueous 2 M HCl, aqueous NaHCO$_3$ solution, saturated, aqueous EDTA solution, and brine, and finally dried over MgSO$_4$. Afterwards, the solvents were removed under vacuum. The resulting solid was dissolved in a small amount of

chloroform and precipitated into 500 mL of methanol to afford the target polymer as a light-green solid. Subsequent Soxhlet extractions were carried out with methanol, acetone, ethyl acetate and chloroform, respectively. After re-precipitation of the chloroform-soluble fraction into methanol, the light-green polymer was obtained with 70% yield (211 mg). ^{1}H NMR (600 MHz, $C_2D_2Cl_4$, 60 °C) δ 8.60–8.41 (m, 2H), 7.86–7.62 (m, 2H), 7.48–6.87 (m, 21H). ^{13}C NMR (150 MHz, $C_2D_2Cl_4$, 60 °C) 143.76, 143.59, 143.43, 143.15, 142.26, 140.72, 140.51, 136.04, 134.62, 132.91, 131.55, 131.53, 131.47, 128.09, 128.01, 127.94, 126.93, 126.86, 126.78, 126.09, 124.46, 119.10, 110.54. M_n 6800, M_w 14 300, and M_w/M_n 2.10 (GPC, PS calibration).

Polymer PCzTPE0.5. A solution of compound **2** (300 mg, 0.457 mmol), compound **3** (200 mg, 0.457 mmol), Ni(COD)$_2$ (629 mg, 2.287 mmol), BPy (357 mg, 2.287 mmol) and COD (247 mg, 2.287 mmol) in 7 mL of THF was reacted under microwave heating at 120 °C for 12 min. The workup procedure was similar to that described for the preparation of PCzTPE. The light-green polymer was obtained with 54% yield (192 mg). ^{1}H NMR (600 MHz, $C_2D_2Cl_4$, 60 °C) δ 8.49 (s, 4H), 8.01–7.66 (m, 4H), 7.58–6.94 (m, 23H), 4.26 (s, 2H), 1.88 (s, 2H), 1.29 (dd, J = 78.1, 27.6 Hz, 10H), 0.92–0.73 (m, 3H). M_n 9200, M_w 23 400, and M_w/M_n 2.54 (GPC, PS calibration).

Results and discussion

Synthesis and characterization

The synthesis routes to the monomers and polymers are depicted in Scheme 1. Monofluoro-TPE **1** was obtained by treating 4-fluorobenzophenone with diphenylmethyl lithium followed by acid-catalyzed dehydration.[46] The N-carbazolyl-TPE derivative **2** was synthesized by catalyst-free N-arylation in a direct nucleophilic substitution of **1** as nonactivated
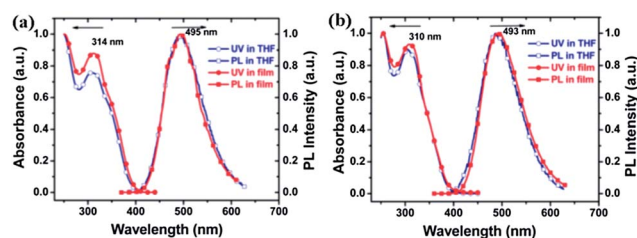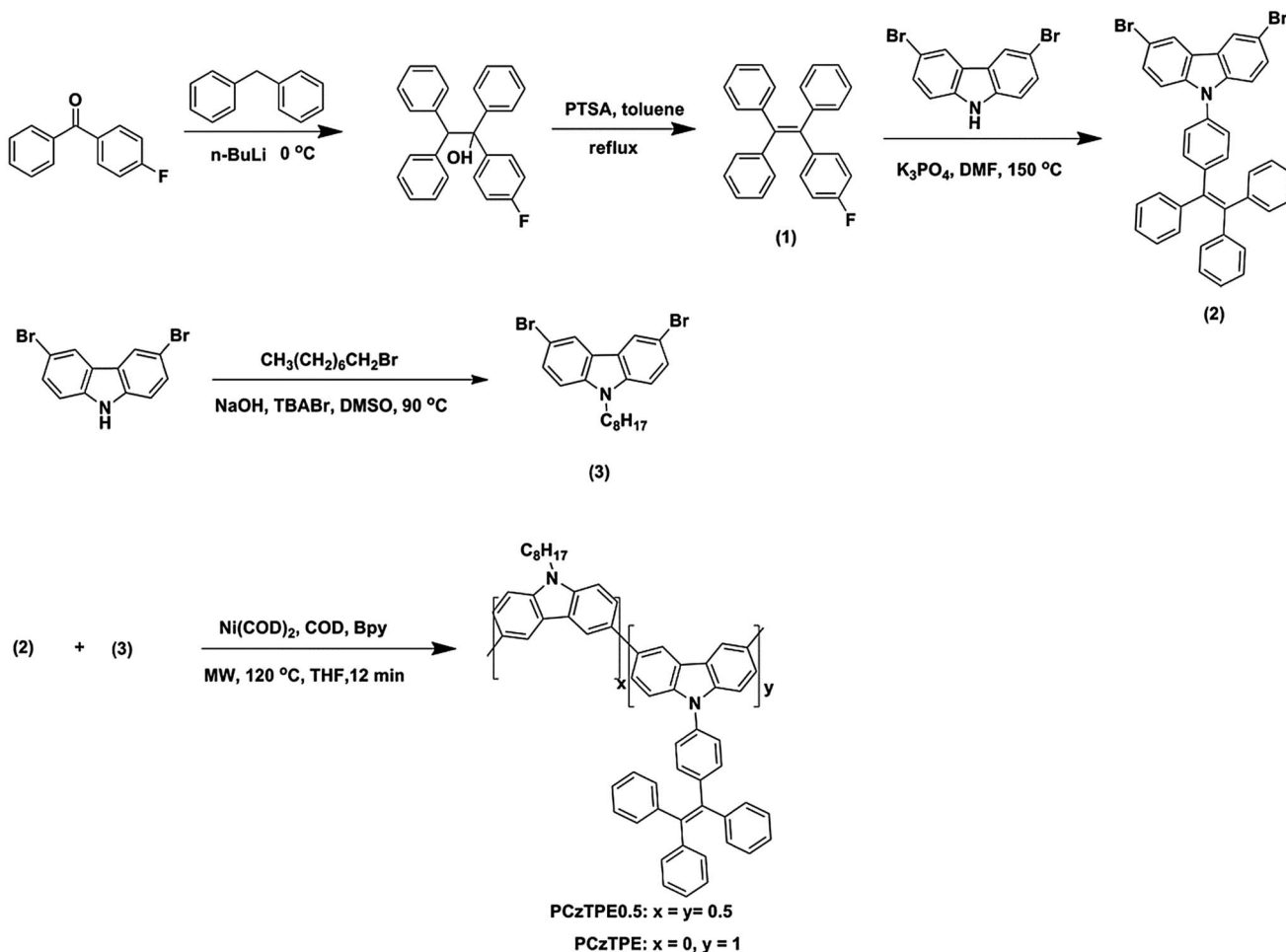

Fig. 1 Absorption and PL spectra (excitation wavelength: 320 nm) of (a) PCzTPE and (b) PCzTPE0.5 in dilute THF solution (10^{-5} M) and as spin-coated films recorded at room temperature.

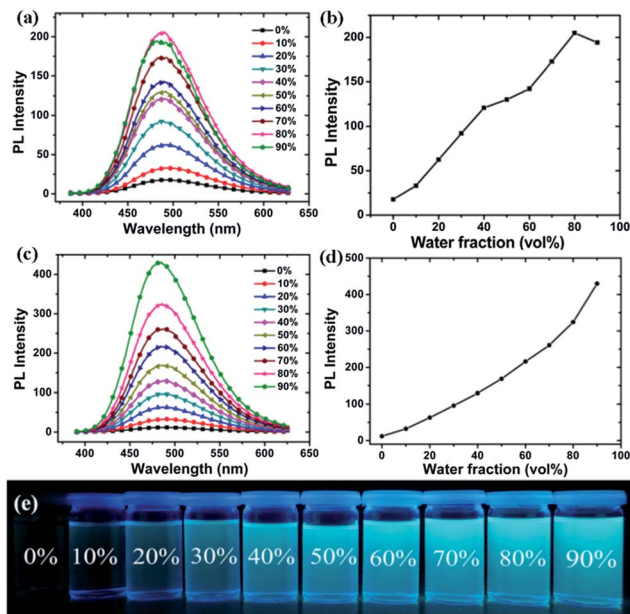Scheme 1 Synthetic procedures for monomer and polymer synthesis.

Fig. 2 PL spectra of (a) PCzTPE and (c) PCzTPE0.5 in THF–water mixtures with different water contents; increase of PL intensity with increasing water fraction for (b) PCzTPE and (d) PCzTPE0.5, respectively (polymer concentration: 10^{-5} M; excitation wavelength: 320 nm); (e) fluorescence images of PCzTPE0.5 in THF–water mixtures with different water contents.

fluorobenzene with 3,6-dibromocarbazole.[47] 3,6-Dibromo-9-octylcarbazole **3** was synthesized according to a reported procedure.[48] All monomers were fully characterized prior to polymerization. Homopolymer PCzTPE and random copolymer PCzTPE0.5 were synthesized from monomers **2** and **3** by Yamamoto-type coupling using Ni(COD)$_2$ as a coupling reagent in a mixture of THF, COD and Bpy under microwave (MW) heating.[49] Following these protocols we could obtain the target conjugated polymers in short reaction times. The chemical structures of the obtained polymers were confirmed by NMR spectroscopy, GPC, thermal analysis and optical spectroscopy.

Thermal properties

The thermal properties of PCzTPE0.5 and PCzTPE were investigated by TGA and DSC. Both polymers exhibit high thermal stability with 5% weight loss occurring at 490 and 430 °C, respectively. In DSC analysis we could not record glass transitions (T_g) up to 300 °C. High thermal stability is important for practical application in solid state sensors.

Photophysical properties

Fig. 1 shows the absorption and PL spectra of PCzTPE and PCzTPE0.5 in THF solution and solid state films. The absorption spectra of PCzTPE and PCzTPE0.5 are very similar, with solid state peak maxima at 314 and 310 nm, respectively. In the PL spectra both polymers exhibit green fluorescence peaking around 495 nm, both in solution and as thin films. This behavior is attributed to the incorporation of the TPE units – they effectively suppress π-stacking in the condensed phase due
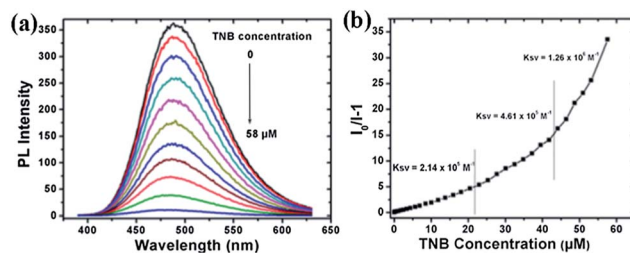

Fig. 3 (a) PL spectra of PCzTPE0.5 in THF–water 1 : 9 containing different amounts of TNB. Polymer concentration: 10^{-5} M; λ_{ex}: 320 nm. (b) Stern–Volmer plots of PL intensity of PCzTPE0.5 $I_0/I - 1$ versus TNB concentration (I = PL intensity, I_0 = PL intensity at a TNB concentration of 0 M).

to the presence of the propeller-shaped TPE side chain. The photoluminescence quantum yields (PLQYs) of PCzTPE and PCzTPE0.5 in diluted THF solution, estimated by using quinine sulfate as the standard, have been determined as 1.1% and 0.8%, respectively. The PLQYs distinctly increase to 20% and 21%, respectively, in solid state films, 18- and 26-fold higher when compared to THF solutions. Evidently, the transition into the condensed state dramatically enhances the PL of the polymers. The AIE properties will be discussed in detail in the following paragraph.

To further investigate the AIE effect with the polymers PCzTPE and PCzTPE0.5, a series of PL spectra in THF–water mixtures with increasing water fraction were recorded. Fig. 2a and c show PL spectra of the polymers in such water–THF mixtures. The PL intensity increases progressively with increasing water fraction for both polymers. Hereby, polymer PCzTPE0.5 showed a more pronounced AIE effect when compared to the PCzTPE copolymer. For PCzTPE the PL intensity is 11 times higher for a water content of 80% when compared to pure THF (Fig. 2b), and for PCzTPE0.5 35 times higher for a water content of 90% (Fig. 2d). As water is a non-solvent for PCzTPE and PCzTPE0.5, both polymers are assumed to form solid-state aggregates in the THF–water mixtures with a high water content, thus exhibiting aggregation-induced PL enhancement: they are AIE active. Caused by the high rotational freedom of the TPE side-chain moieties in solution, a high internal conversion rate results in weak emission. Within the aggregated (solid) state, however, the rotation of phenyl rings of
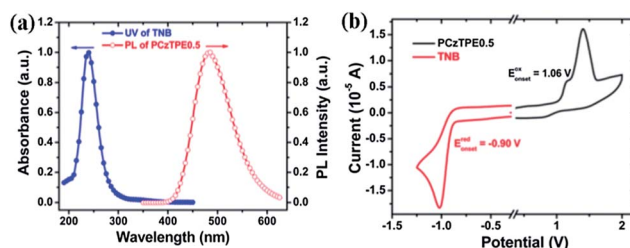

Fig. 4 (a) Normalized absorption spectrum of TNB and the PL spectrum of PCzTPE0.5 nanoaggregates in THF–water 1 : 9. (b) CV plots of PCzTPE0.5 in dichloromethane (oxidative scan) and TNB in acetonitrile (reductive scan), for the experimental conditions please see the Characterization of materials chapter.
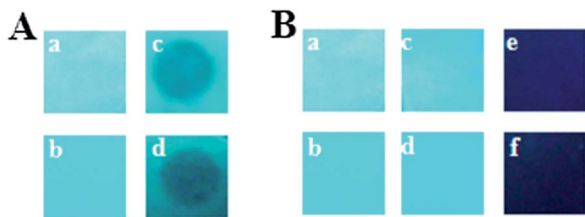
Fig. 5 Paper strip tests. (A) Vapor-mode detection of TNB: test strips before (a and b) and after (c and d) placing the strips on top of a glass vial containing solid TNB for 5 min; for PCzTPE (a and c) or PCzTPE0.5 (b and d), respectively. (B) Solution-mode detection of TNB: test strips before (a and b) and after dipping the strips into pure THF (c and d) and into a 10^{-4} M TNB solution in THF (e and f); for PCzTPE (a, c and e) or PCzTPE0.5 (b, d and f), respectively.

the TPE units is strongly restricted, thus blocking non-radiative deactivation channels and leading to the AIE effect. Poly(N-octyl-3,6-carbazole) without TPE side groups that was generated as a reference did not show any AIE activity.

Explosive detection

PCzTPE and PCzTPE0.5 should show good electron-donor ability due to their electron-rich polycarbazole backbones. Moreover, their twisted 3D-structure should create effective pathways for interchain exciton diffusion leading to amplified PL quenching properties. Based on the above presented results we started our investigations with aggregated PCzTPE0.5 in THF–water 1 : 9 (polymer concentration 10 μM). 1,3,5-Trinitrobenzene (TNB) was chosen as a prototypical nitroaromatic analyte. As shown in Fig. 3a, the PL intensity of PCzTPE0.5 in THF–water 1 : 9 decreases progressively during addition of TNB, without changing the PL peak position, suggesting that different emissive species are not formed. The onset of PL quenching is found for addition of 50 nM TNB, low enough for the detection of submillimolar TNB concentrations. For a TNB concentration of 58 μM, the PL of the dispersed polymer nanoaggregates is fully quenched. The quenching response was analyzed by fitting the data with the Stern–Volmer equation, as depicted in Fig. 3b. For TNB concentrations below 21 μM, the Stern–Volmer plot is linear with a quenching constant of 2.14×10^5 M^{-1}. For a higher analyte concentration, the curve bends upward, thus demonstrating an amplified quenching.[50] The quenching constant reaches ca. 1.26×10^6 M^{-1} between TNB concentrations of 43 μM and 58 μM. This amplified quenching is attributed to the twisted 3D topology of the polymer chains in the nanoaggregates, leading to the formation of an increased number of quenching sites that can interact with TNB molecules and/or to an improved exciton diffusion to quenching sites.[51]

Fig. 4a shows that there is no spectral overlap between the absorption spectrum of TNB and the PL spectrum of PCzTPE0.5 (as a prerequisite for Förster-type energy transfer) thus indicating that the main quenching mechanism for TNB addition should be an excited state charge transfer between the excited state of the host and the ground state of the TNB quencher. The occurrence of charge transfer was further confirmed by cyclic voltammetry (Fig. 4b). The HOMO (highest occupied molecular

orbital) level of PCzTPE0.5 was estimated to be ca. −5.1 eV, and the LUMO (lowest unoccupied molecular orbital) level of TNB to be ca. −3.1 eV. Considering an optical bandgap (E_g) of 3.2 eV for PCzTPE0.5 from the onset of its UV/vis absorption band, the LUMO level of PCzTPE0.5 is calculated to be ca. −1.9 eV. Therefore, the LUMO energy (−1.9 eV) of PCzTPE0.5 allows for an excited state electron transfer to the lower-lying LUMO level of TNB with a LUMO–LUMO offset of ca. 1.2 eV.[52]

For practical explosive detection, the availability of solid state sensor devices is of primary importance. Towards this goal, we prepared test strips by dip-coating Whatman filter paper into solutions of PCzTPE and PCzTPE0.5 in THF (10^{-4} M) followed by drying the strips in an air stream. First, for vapor-mode tests, we placed the fluorescent paper strips on top of a glass vial containing solid TNB for 5 min at room temperature. In this way, a circular area of the strip was exposed to TNB vapor. Within the exposed area the PL of the polymers was obviously distinctly quenched (Fig. 5A). Second, for solution tests, the test strips containing both polymers were dipped into pure THF (as reference) and a solution of TNB in THF (10^{-4} M). As shown in Fig. 5B, the fluorescence of the strips was quenched completely after contact with the TNB solution for both polymers. The reference strips dipped into pure THF did not show significant PL quenching thus demonstrating that the majority of the polymers remain adsorbed at the test strips. These first promising results demonstrate the potential of our new polycarbazole-type polymers PCzTPE0.5 and PCzTPE for the fabrication of solid state sensors for nitroaromatic explosives with sufficient sensitivity.

Conclusions

In summary, two poly(3,6-carbazole)s with AIE-active tetraphenylethylene (TPE) side chains have been successfully synthesized. The polymers combine the electron-deficient character of the polycarbazole backbone and the AIE behavior of the TPE containing side chains. So, both polymers PCzTPE and PCzTPE0.5 show distinct AIE properties. For sensing of nitroaromatic explosives PL quenching experiments were carried out. Aggregated PCzTPE0.5 shows amplified PL quenching during trinitrobenzene (TNB) addition in THF–water mixtures (1 : 9, v/v) with a maximum Stern–Volmer quenching constant of 1.26×10^6 M^{-1}. Solid-state paper strips with deposits of both polymers show TNB-induced PL quenching, both towards TNB vapor or TNB solution thus demonstrating promising practical application potential in solid state sensors for nitroaromatic explosives. Other possible applications (e.g. in OLEDs) are currently tested.

Acknowledgements

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