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# Unlocking the chemical environment of nitrogen in perovskite-type oxides†

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Nitrogen (N) doping of perovskite-type oxides is an effective method for enhancing their photocatalytic performance. Quantitative and qualitative analyses of the doped N species are essential for a deeper understanding of the catalytic activity enhancement mechanism. However, examining the N environment in perovskite-type oxides, particularly in the bulk, using conventional analytical techniques, such as X-ray photoelectron spectroscopy (XPS), is challenging. In this study, we propose a new analytical technique, advanced temperature-programmed desorption (TPD) up to 1600 °C, to complement the conventional methods. TPD can quantify all N species in bulk oxides. Moreover, it facilitates chemical speciation of N environments, such as substitutional and interstitial N species. This is verified by XPS, CHN elemental analysis, X-ray absorption spectroscopy, and *in situ* diffuse reflectance infrared Fourier-transform spectroscopy. This study demonstrates the feasibility of advanced TPD as a new analytical method that offers comprehensive information on the N species within N-doped oxide materials at the bulk level.

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## Introduction

Perovskite-type oxides are functional materials with excellent physical and chemical properties.<sup>1–4</sup> In particular, two-dimensional perovskite-type oxides have recently gained significant attention for solar energy conversion and environmental remediation.<sup>5–8</sup> Perovskite-type La<sub>2</sub>Ti<sub>2</sub>O<sub>7</sub> (LTO) nanoparticles with a layered structure has been regarded as a promising photocatalyst for organic decomposition, water splitting, and CO<sub>2</sub> reduction reactions owing to its high charge carrier mobility, high stability, low cost, and low toxicity.<sup>6,9–12</sup> However, LTO photocatalysts are inefficient at utilizing visible light because of their wide bandgap of approximately 3.8 eV.<sup>13,14</sup> Therefore, N-doping is an effective strategy for narrowing the bandgap.<sup>15,16</sup> Generally, the doped N species in LTO belong to two chemical states: (i) substitutional N for oxygen atoms in the

lattice (Ti–N bond) or (ii) interstitial N bonding to oxygen in the lattice structure (Ti–O–N bond).<sup>17–19</sup> These introduction modes of N species have a significant impact on the resulting band structure. Theoretical studies have revealed a correlation between the N environment and photocatalytic activity.<sup>12,20</sup> For substituted N species, mixing of the doped N 2p and O 2p orbitals creates continuous bands above the valence band maximum, thus narrowing the band gap. For interstitial N, the π\* NO levels originating from the Ti–O–N bond were in the high-energy region of the band gap, enabling visible light adsorption. However, these π\* NO levels also act as recombination centres for charge carriers and may reduce the photocatalytic activity. It has been experimentally shown that an appropriate ratio of substitutional N to the interstitial N species results in high photocatalytic activity.<sup>16</sup> Therefore, precisely tuned N-doping is essential for targeted band engineering to develop visible-light-responsive catalysts.

Precise N-doping, first of all, requires a reliable analytical technique that can distinguish and quantify the forms of doped N species. In most studies, X-ray photoelectron spectroscopy (XPS) has been used to analyse the N species in LTO and examine the electronic states.<sup>9,15,16,21</sup> However, there is an issue with the inconsistent peak assignments of N-doped LTO in the N 1s XPS spectra. Most reports attribute the peak at approximately 399 eV to the Ti–O–N bond,<sup>19,22–27</sup> whereas some attribute it to the Ti–N bond.<sup>28,29</sup> In addition, XPS measurements can provide only the surface environment; however, to fully understand the doping states of the N species, bulk information is required. CHN elemental analysis can be used to

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† Electronic supplementary information (ESI) available: PXRD patterns, SEM images, N 1s and C 1s XPS spectra, and TPD results of LTO\_TEA and LTO\_N, MS and *in situ* DRIFTS spectra of LTO\_TEA, TPD result of TiN, the calibration curves for TPD, and a schematic illustration of the advanced TPD. See DOI: <https://doi.org/10.1039/d4sc01850h>



determine the total N content in the bulk but is not suitable for chemical speciation. Therefore, another new technique that can accurately perform both quantitative and qualitative analyses of bulk N species is required.

Temperature-programmed desorption (TPD) is an important thermal analytical method, in which a sample is heated at a constant heating rate, and the desorbed gases from the sample are detected using a mass spectrometer or high-speed gas chromatography.<sup>30–33</sup> For example, when the TPD method is applied to carbon materials, each oxygen functional group can be quantified from the information on the desorbed gas species and their desorption temperature. However, the total determination of hydrogen in carbon materials is not possible with commercial equipment because of their low sensitivity under a He flow.<sup>34,35</sup> Our group developed an advanced vacuum TPD device with an induction heating system capable of reaching temperatures up to 1800 °C.<sup>36–40</sup> This technology enables total hydrogen quantification in carbon.<sup>36–39</sup> More recently, the maximum temperature of TPD has been extended to 2100 °C, allowing qualitative and quantitative analysis of N species in bulk carbon materials, with high sensitivity at the 10 ppm level.<sup>41</sup> High sensitivity has been achieved by thorough insulation of the sample holder to minimize heat transfer at very high temperatures, which suppresses degassing from the TPD equipment components.<sup>41</sup> Owing to these features, advanced TPD paves the way for new frontiers in carbon-edge site chemistry, leading to the development of three-dimensional graphene-based materials.<sup>42–48</sup> Here, it is conceivable that the advanced TPD method could be useful not only for carbon materials but also for oxides doped with light elements. An essential feature of the TPD method is that it is a bulk analytical method, complementing the surface-sensitive XPS techniques.

In this study, we propose high-temperature vacuum TPD as a promising evaluation approach for N species in oxides. The measurement targets used in this study were perovskite-type N-doped LTO nanoparticles.<sup>16</sup> We conducted a combined study of XPS, X-ray absorption spectroscopy (XAS), CHN elemental analysis, and *in situ* diffuse reflectance infrared Fourier-transform spectroscopy (DRIFTS). Our findings revealed that advanced TPD can quantify the total N species in bulk materials; more importantly, the features of the TPD spectra reflect the N environment in perovskite-type oxides, such as substitutional and interstitial N.

## Experimental

### Sample preparation

**Synthesis of La<sub>2</sub>Ti<sub>2</sub>O<sub>7</sub> nanoparticles (LTO\_TEA).** La<sub>2</sub>Ti<sub>2</sub>O<sub>7</sub> nanoparticles were synthesized according to the procedure described in a previous report.<sup>16</sup> Titanium(IV) bis(ammonium lactate)dihydroxide solution (C<sub>6</sub>H<sub>18</sub>N<sub>2</sub>O<sub>8</sub>Ti, 50 wt% in H<sub>2</sub>O, Aldrich, denoted as TAL) was used as a titanium source. La(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O (>99%, Wako Pure Chemicals) was used as the lanthanum source. Triethanolamine (TEA, min 98.0%, Kanto Chemical, denoted as TEA) was employed as an additive to minimize particle size. Initially, 10 mL of deionized water was

mixed with 0.96 mL of TAL and 2 mmol of La(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O, and the solution was stirred for 20 min. Subsequently, 10 mL of 1.5 M NaOH aqueous solution and 8 mmol of TEA were gently added. After 30 min of homogenization, the mixture was put into a 100 mL Teflon-lined stainless-steel autoclave for hydrothermal treatment at 240 °C for one day. Subsequently, the mixture was allowed to cool to ambient temperature. Following treatment, the white precipitate was washed with deionized water and centrifuged multiple times for complete separation. The product was dried for an entire night at 60 °C under low pressure. The obtained white powder was denoted as LTO\_TEA.

**Synthesis of nitrogen-doped La<sub>2</sub>Ti<sub>2</sub>O<sub>7</sub> nanoparticles (LTO\_N).** The N-doped La<sub>2</sub>Ti<sub>2</sub>O<sub>7</sub> nanoparticles (denoted as LTO\_N) were produced by heating LTO\_TEA in a tube furnace to 650 °C under an NH<sub>3</sub> flow at a rate of 50 mL min<sup>-1</sup>.<sup>16</sup> The La<sub>2</sub>Ti<sub>2</sub>O<sub>7</sub> powder was placed in an alumina boat at the centre of a tube furnace containing a gas flow meter to regulate the flow rate. After 3 h, the samples were allowed to cool naturally to room temperature in an NH<sub>3</sub> atmosphere.

**Characterization.** X-ray photoelectron spectroscopy (XPS) was performed with a ULVAC PHI 5600 XPS system using the Al K $\alpha$  line (1486.6 eV). Ar<sup>+</sup> sputtering was used to analyse the depth profiles of the samples. The binding energy before the Ar<sup>+</sup> sputtering was calibrated using the C 1s photoelectronic peak at 284.6 eV. It is noteworthy that in each XPS spectrum after Ar<sup>+</sup> sputtering, the correction was performed in the same manner as before Ar<sup>+</sup> sputtering. This is because the C 1s peak became negligible after the Ar<sup>+</sup> sputtering process. To determine the N content by XPS, the samples were assumed to be composed of C, N, O, Ti, and La. Powder X-ray diffraction (PXRD) patterns were collected using a Rigaku MiniFlexII between 10 and 60° (2 $\theta$ ) with a 0.01° step (conditions: Cu K $\alpha$  radiation, 45 kV, 15 mA, and 1.5418 Å). Scanning electron microscopy (SEM) images were obtained on a FE-SEM using a Hitachi S4800. Thermal analysis for comparison with TPD was performed using thermogravimetry-mass spectrometry (TG-MS). The measurements were conducted on a Netzsch STA 449 Jupiter spectrometer by heating the samples from 60 to 1600 °C at a heating rate of 10 °C min<sup>-1</sup> under a He flow of 150 mL min<sup>-1</sup>. The emitted gas was analysed using a quadrupole mass spectrometer (JEOL JMS-G1500GC). Because of the slight weight changes, the TG curves did not provide useful information and were not examined in this study. The instrument was placed in a box filled with Ar gas during the measurement to minimize air contamination in the sample chamber.<sup>49–51</sup> CHN elemental analysis was performed using a J-Science Lab JM10 CHN Analyzer MICRO CORDER. The samples were combusted under a flow of 80% He and 20% O<sub>2</sub>, and the gases produced were converted into CO<sub>2</sub>, H<sub>2</sub>O, and N<sub>2</sub> to determine the amounts of C, H, and N, respectively. N K-edge X-ray absorption spectroscopy (XAS) measurements were performed in total electron yield (TEY) and partial fluorescence yield (PFY) modes using synchrotron radiation at beamline BL12 at the SAGA Light Source, Japan. The energy was calibrated by comparing the TiN results with those in the literature.<sup>52</sup> *In situ* diffuse reflectance infrared Fourier-transform spectroscopy (*in situ* DRIFTS) experiments were performed using an S.T. Japan Inc. Heat



Chamber Type-1000 °C/DiffusIR mounted on a Shimadzu IRSpirit-X under vacuum.

### TPD analysis

**High-temperature vacuum TPD system.** To study the desorption behaviour of the N species in the oxides, TPD measurements were performed using an advanced TPD system developed by our group.<sup>36,41</sup> The TPD instrument had two sections: one for heating and the other for analysis (Fig. S1a†). In the heating section, the sample holder was heated using a high-frequency induction heater, and its temperature was monitored using a radiation thermometer. Graphite sample holders were used to measure the carbon materials. However, because graphite and oxides react at high temperatures, a tungsten sample holder was used in this study (Fig. S1b†). In the analysis section, a quadrupole mass spectrometer (MPH-100 M, Inficon Co., Ltd.) was used to analyse the gas species desorbed from the sample. Prior to each TPD measurement, an empty tungsten sample container was heated at 1600 °C for 1 h under vacuum to minimize the contamination in measurements. After placing approximately 2 mg of the sample in the sample holder under air, the chamber was evacuated below  $2 \times 10^{-5}$  Pa, and then, the sample was heated at a rate of  $10^\circ \text{C min}^{-1}$  from room temperature to 1600 °C for the TPD measurement. The detailed methods for the quantitative evaluation are described below.

**Calibration curve preparation.** Gases desorbed during heating ( $\text{N}_2$ ,  $\text{NO}$ ,  $\text{NH}_3$ ,  $\text{HCN}$ ,  $\text{H}_2$ ,  $\text{H}_2\text{O}$ ,  $\text{CO}$ , and  $\text{CO}_2$ ) were quantified using the developed calibration curves to convert the peak areas into gas quantities.<sup>41</sup> The calibration curves for  $\text{N}_2$ ,  $\text{NH}_3$ ,  $\text{H}_2$ ,  $\text{CO}$ ,  $\text{NO}$ , and  $\text{CO}_2$  were prepared by introducing the gases from gas cylinders into the TPD chamber. To determine the exact amount of gas introduced, a gas reservoir tank of known volume (Fig. S1a†) was filled with each gas at a pressure of 10–35 Pa. Subsequently, the pressure change was monitored before and after the introduction of the gas into the TPD chamber. A calibration curve for  $\text{H}_2\text{O}$  was prepared by introducing pure water vapor produced from deionized water, which was degassed using the freeze–pump–thaw cycling method in an Ar atmosphere.  $\text{HCN}$  gas was produced by adding  $\text{NaCN}$  (FUJIFILM Wako Pure Chemicals Co., Ltd.) to phosphoric acid, and a mixture of  $\text{HCN}$  and air was introduced into the TPD chamber. The slopes of the calibration curves of each gas species [ $\mu\text{mol A}^{-1} \text{s}^{-1}$ ] are listed in Table S1.†

**Quantification of  $\text{H}_2$ ,  $\text{H}_2\text{O}$ ,  $\text{CO}_2$ ,  $\text{NO}$ , and  $\text{HCN}$ .** The desorption rate ( $V_{\text{des}}$  [ $\mu\text{mol g}^{-1} \text{s}^{-1}$ ]) was calculated as follows:

$$V_{\text{des}} = \frac{I_{\text{des}} a}{W_{\text{sample}}} \quad (1)$$

where  $I_{\text{des}}$ ,  $a$ , and  $W_{\text{sample}}$  correspond to the intensity of the corresponding mass number of the desorbed gas species [A] (Fig. S2a and b†), the slope of the prepared calibration curve [ $\mu\text{mol A}^{-1} \text{s}^{-1}$ ] (Table S1†), and the sample weight [g].

**$\text{NH}_3$  desorption.** The total  $m/z = 17$  profile included contributions from both  $\text{NH}_3$  and  $\text{OH}$ , as both had a mass number of  $m/z = 17$ .<sup>41</sup> To extract the  $\text{NH}_3$  component from the TPD profile

for  $m/z = 17$ , an  $\text{OH}$  profile was generated using a  $\text{H}_2\text{O}$  ( $m/z = 18$ ) profile. The profile was generated considering that the intensity of  $\text{OH}$  is 0.26 times that of  $\text{H}_2\text{O}$  in the calibration curve. The  $\text{NH}_3$  profile was obtained by subtracting the  $\text{OH}$  contribution from the total  $m/z = 17$  profile (Fig. S2c and d†).

**$\text{CO}$  and  $\text{N}_2$  desorption.** The total  $m/z = 28$  profile contained contributions from  $\text{CO}$  and  $\text{N}_2$ . To separate the  $\text{CO}$  and  $\text{N}_2$  traces in the  $m/z = 28$  profile, an  $\text{N}_2$  profile was created using the  $\text{N}$  ( $m/z = 14$ ) profile. The  $\text{N}$  fragment ( $m/z = 14$ ) was formed not only from  $\text{N}_2$  but also from  $\text{HCN}$ ,  $\text{NH}_3$ , and  $\text{NO}$ . To extract the  $\text{N}$  fragment profile from  $\text{N}_2$ ,  $\text{N}$  profiles from  $\text{HCN}$ ,  $\text{NH}_3$ , and  $\text{NO}$  were generated using the gas profiles of  $m/z = 27$ , 17, and 30, respectively. The profiles were generated considering that the intensity of  $\text{N}$  is 0.02 times that of  $\text{HCN}$ , 0.03 times that of  $\text{NH}_3$ , and 0.07 times that of  $\text{NO}$ . The  $\text{N}$  fragment from the  $\text{N}_2$  profile was then obtained by subtracting the  $\text{HCN}$ ,  $\text{NH}_3$ , and  $\text{NO}$  contributions from the total  $m/z = 14$  profile (Fig. S2e and f†). The  $\text{N}_2$  profile was created based on the fact that the intensity of  $\text{N}_2$  was 16 times that of  $\text{N}$ . Finally, the  $\text{CO}$  profile was obtained by subtracting the  $\text{N}_2$  contribution from the total  $m/z = 28$  profile (Fig. S2g and h†).

## Results and discussion

### TPD analysis of N species in $\text{La}_2\text{Ti}_2\text{O}_7$

For the TPD measurements, perovskite-type N-doped LTO nanoparticles were synthesized as target materials according to a previous report.<sup>16</sup> The synthesis procedure is illustrated in

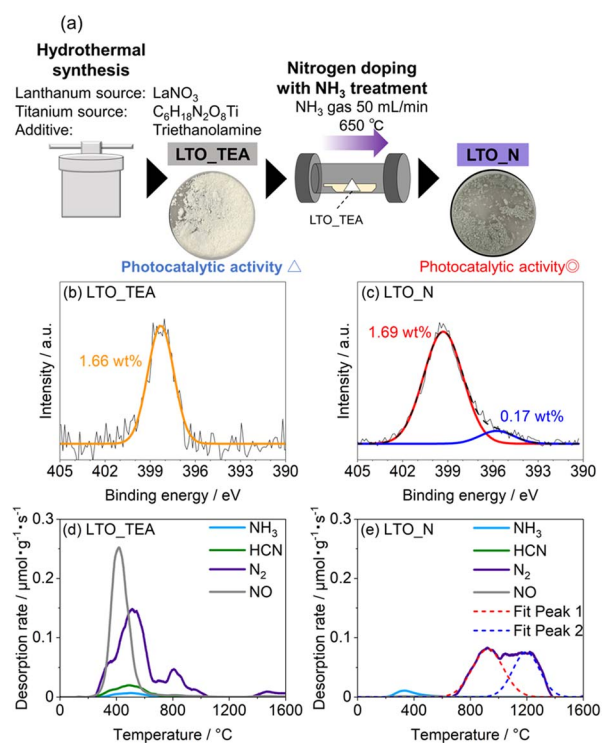


Fig. 1 (a) Schematic of the synthesis procedure of LTO\_TEA and LTO\_N. N 1s XPS spectra of (b) LTO\_TEA and (c) LTO\_N. TPD profiles of (d) LTO\_TEA and (e) LTO\_N for N-containing gases.



Fig. 1a (see the Experimental section for details). Briefly, LTO nanoparticles (LTO\_TEA) were synthesized through a hydrothermal method using lanthanum and titanium sources along with triethanolamine (TEA) as an additive. LTO\_TEA nanoparticles were white in colour, containing trace amounts of N species derived from TEA on their surface. The prepared LTO\_TEA nanoparticles were heated at 650 °C under an NH<sub>3</sub> flow to produce grey LTO nanoparticles with N (LTO\_N). Nitrogen in the nanoparticles, originating from NH<sub>3</sub>, was introduced into the matrices by diffusion from the surface to the bulk. A previous study demonstrated that LTO\_N has higher photocatalytic activity than LTO\_TEA under visible light irradiation because of its narrower band gap caused by N doping *via* NH<sub>3</sub> treatment.<sup>16</sup> Note that the PXRD patterns and XPS spectra of LTO\_TEA and LTO\_N agreed well with the previous study, confirming synthetic reproducibility (see Fig. 1b, c, and S3†).<sup>16</sup>

N 1s XPS measurements were performed to gain insight into the chemical states of the N species. As shown in Fig. 1b, LTO\_TEA exhibited a peak at 398.3 eV, corresponding to 1.66 wt% N species even before the NH<sub>3</sub> treatment. The presence of a trace amount of N is ascribed to contaminants derived from TEA added during the hydrothermal process. LTO\_N, obtained after NH<sub>3</sub> treatment, exhibited a peak at approximately 399.3 eV (red), with a small shoulder peak at approximately 395.8 eV (blue) (Fig. 1c). The peak at 399.3 eV can be assigned to Ti–O–N bonds, that is, interstitial N bonding with oxygen in the lattice, based on previous literature.<sup>26,27</sup> The peak at 395.8 eV is attributed to Ti–N bonds, that is, substitutional N for oxygen in the framework.<sup>6,19,26,53</sup> However, determining the N environment solely from the XPS peak assignments is difficult. First, as mentioned in the Introduction section, various assertions have been presented in the literature concerning the attribution of peaks to particular chemical species. Second, several different N species often have similar binding energies. The major peaks in LTO\_TEA and LTO\_N were observed at

approximately 399 eV (see Fig. S4†). However, these samples, which were synthesized through different processes, may contain different types of N species (discussed in more detail in the next section). Therefore, relying solely on XPS measurements can lead to inappropriate conclusions regarding the N states of the oxide.

In this study, we developed an advanced TPD method for bulk analysis of N-doped metal oxides. We previously developed a highly sensitive TPD system for analysing carbon-based materials (Fig. S1a†). However, graphite sample holders, which have been used for carbon samples, react with metal oxides at high temperatures, even in a vacuum, to produce CO gas.<sup>54,55</sup> Thus, we developed a new system that allows the use of tungsten sample holders with high melting and boiling points because they do not react with metal oxides even at high temperatures (Fig. S1b†).<sup>56</sup> Minimal gas desorption was confirmed at temperatures up to 1600 °C in the blank test using a tungsten sample holder (Fig. S5†). Fig. 1d shows the TPD profiles of the N-containing gases of LTO\_TEA (for analysis details, see the Experimental section and Fig. S2 and S6†). The main desorbed gas species were NO and N<sub>2</sub>, with the desorption temperatures ranging from 300 to 600 °C and 300 to 1000 °C, respectively. In contrast, in the desorption profile of LTO\_N (Fig. 1e), the desorbed gas at 700–1400 °C was primarily N<sub>2</sub>. These results indicate that the desorption pattern depends on the environment of the N species, suggesting the potential use of TPD for the identification of N species. This has been described in detail in the next section. The quantitative accuracy of TPD was verified by comparison with other analytical methods, and the corresponding data are presented in Table 1. The N contents determined by TPD align closely with the CHN results, providing clear evidence that the N species were fully detected as gas species by TPD. In contrast, XPS revealed larger quantitative values for both samples. This difference is likely due to the surface segregation of N species in the samples, given

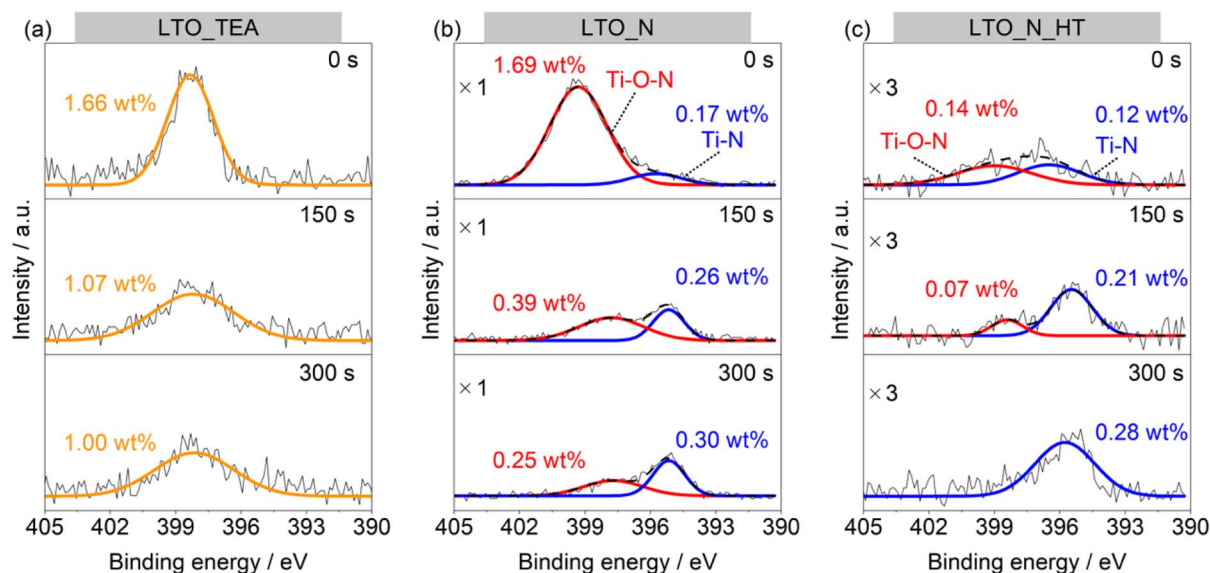


Fig. 2 N 1s XPS spectra of (a) LTO\_TEA, (b) LTO\_N, and (c) LTO\_N\_HT before and after Ar<sup>+</sup> sputtering for 150 s or 300 s.



**Table 1** Nitrogen contents in LTO\_TEA and LTO\_N determined by different analytical methods

Sample	TPD/wt%	CHN/wt%	XPS/wt%
LTO_TEA	1.19	1.13	1.66
LTO_N	0.68	0.72	1.86

that XPS is surface-sensitive. This segregation is reasonable because N-containing organic contaminants in LTO\_TEA are expected to be anchored to the surface. Moreover, the N source of LTO\_N is dissolved in the solid and diffuses from the surface to the bulk during synthesis. The TPD results align well with those obtained from the CHN analysis, in contrast to XPS. This strongly indicates that TPD is a bulk analytical method for oxides. The desorption of all N species at temperatures up to 1600 °C can be verified by the N 1s XPS spectra (Fig. S7†) and CHN elemental analysis results (Table S2†) of post-TPD measurements. Note that CHN elemental analysis overestimates the amount of N species below 0.1 wt%, always showing approximately 0.1 wt%.<sup>41</sup>

Commercial instruments generally detect desorbed gases under a He flow rather than under vacuum. Fig. S8† shows the results of the desorbed gas analysis of LTO\_TEA using a widely used thermogravimetry-mass spectrometry (TG-MS) system. The peak of  $m/z = 30$  (NO) was observed around 350 °C; however, the signal for  $m/z = 28$ , corresponding to N<sub>2</sub>, continuously decreased, and no clear desorption peak was observed. These results are significantly different from the advanced vacuum TPD results (see Fig. 1d). The commercial system could not detect trace N<sub>2</sub> gas desorption from the sample owing to air contamination in the chamber, despite minimizing the air contamination (see also the Experimental section). This highlights the superiority of the advanced TPD method using induction heating under vacuum for N analysis.

### Desorption pattern assignments for TPD

To establish a qualitative analytical method, examining the types of N species in the TPD profiles that are responsible for NO and N<sub>2</sub> gas desorption from LTO\_TEA and LTO\_N is important. The TPD profile of LTO\_TEA exhibited NO and N<sub>2</sub> desorption at low temperatures, along with CO, CO<sub>2</sub>, and H<sub>2</sub>O desorption (Fig. 1d and S6a†). The overlap of N-containing gas evolution with those typically found during the thermal decomposition of organic substances strongly suggests that NO and N<sub>2</sub> were generated by the decomposition of organic substances derived from the TEA used in the hydrothermal reaction. It has been reported that N species in organic compounds exhibit an N 1s peak at a binding energy of 398–399 eV.<sup>57–59</sup> Moreover, the electrostatic attraction between the N of a TEA molecule and the surface ≡TiO<sup>−</sup> group (N–O bond) also exhibits an N 1s peak at a binding energy of 398 eV.<sup>60</sup> These observations are consistent with the peak position of LTO\_TEA, suggesting the presence of N–O and/or C–N bonds (Fig. 1b). In contrast, N<sub>2</sub> desorption from LTO\_N at temperatures above 800 °C was not accompanied by the desorption of other gases (see

Fig. 1e and S6b†). Only CO<sub>2</sub> desorption, presumably from residual organic substances, was observed at 200–800 °C. Thus, the N<sub>2</sub> desorption in a high-temperature region can be ascribed to the strongly bound N species in the oxides. It is noteworthy that both LTO\_TEA and LTO\_N exhibited peaks at similar binding energies in the N 1s XPS spectra (see Fig. S4†); therefore, discussing the N environment based on XPS alone could not provide accurate inferences. The combination of TPD and XPS facilitates a more reliable chemical structure analysis.

As mentioned previously, there are two main doped forms of N species in perovskite-type oxides: substitutional (Ti–N bonds) and interstitial (Ti–O–N bonds). If TPD can distinguish between these N species, it could be a useful analytical tool. Thus, we attempted to deconvolute the N<sub>2</sub> emission pattern (Fig. 1e) using the depth profiles of the N 1s and C 1s XPS spectra obtained by Ar<sup>+</sup> sputtering (Fig. 2 and S9†). LTO\_TEA exhibited a peak at almost the same position even after Ar<sup>+</sup> sputtering in both the N 1s and C 1s XPS profiles. This indicates that similar species (N-containing organic substances) are present from the surface to the interlayer, considering that LTO is a layered perovskite-type oxide. On the other hand, in the case of LTO\_N, the dominant N species at the surface and the interior were different. The concentration of the Ti–O–N species decreased significantly toward the interior of LTO\_N compared to the surface, while the Ti–N species, initially present in low concentration at the surface, became the predominant species in the interior. The amount of carbon inside LTO\_N was negligible (Fig. S9b†), supporting N bonding within the oxide matrices, unlike LTO\_TEA. Further analysis focused on the fact that the N<sub>2</sub> desorption pattern of LTO\_N can deconvolve into two peaks, one at around 900 °C and the other at around 1200 °C (Fig. 1e). To determine the TPD peak assignments of N species (Ti–O–N and Ti–N), a heat-treated LTO\_N sample at 1100 °C was prepared (denoted as LTO\_N-HT). At this temperature, the component of Fit Peak 1 was almost desorbed (Fig. 1e). In other words, LTO\_N-HT may contain N species corresponding to Fit Peak 2. As shown in the N 1s XPS spectra (Fig. 2c), LTO\_N-HT retains the Ti–O–N species at the surface, albeit in a reduced quantity of 0.14 wt% compared to the LTO\_N sample prior to heat treatment (Fig. 2b), with their presence nearly negligible in the interior. In contrast, the amounts of the Ti–N species retained by LTO\_N-HT at the surface and in the bulk were 0.12 wt% and 0.28 wt%, respectively, indicating no significant change after the heat treatment. These results suggest that the Ti–O–N species decomposed during the heat treatments at temperatures up to 1100 °C, while stronger Ti–N bonds were thermally stable, even at 1100 °C. To verify the thermal stability of the Ti–N species, we performed TPD measurements on titanium nitride powder (Wako Pure Chemicals, TiN), as shown in Fig. S10.† Interestingly, N<sub>2</sub> desorption from the TiN matrices can be observed only above 1050 °C, which is in good agreement with the above discussion. The PXRD patterns and SEM images of the LTO\_TEA and LTO\_N samples after heat treatment (Fig. S3†) show no significant structural or morphological changes up to 1100 °C. On the other hand, after heat treatment at 1600 °C, the crystallinity of La<sub>2</sub>Ti<sub>2</sub>O<sub>7</sub> increased and the particle size became larger. Thus, interstitial N species desorb



below 1100 °C without structural or morphological change, while substitutional N species desorb above 1100 °C with a crystallinity increase of  $\text{La}_2\text{Ti}_2\text{O}_7$ . The high thermal stability of  $\text{La}_2\text{Ti}_2\text{O}_7$  up to 1100 °C allows for the sequential desorption of interstitial and substitutional N species, rather than simultaneous desorption and thus enables qualitative analysis by the TPD method.

To further study the N environment at the surface and in the bulk, synchrotron-based XAS analysis was conducted in the total electron yield (TEY) and partial fluorescence yield (PFY) modes. TEY is a surface-sensitive analysis mode, in contrast to PFY, which reveals the internal structure. Thus, XAS can distinguish between the external and internal chemical structures of samples without the  $\text{Ar}^+$  sputtering process, unlike XPS. Fig. 3a and b show the N K-edge X-ray absorption near-edge (XANES) spectra in the TEY and PFY modes, respectively. LTO\_TEA exhibited a peak at 405 eV in both the TEY and PFY modes, which can be attributed to the N–O bond.<sup>61,62</sup> The presence of this peak at 405 eV explains the large amount of NO desorption observed in TPD (Fig. 1d) and is consistent with the XPS depth profiles, indicating the presence of a N–O bond. The N–O bond peak at 405 eV was also observed in the TEY XANES spectrum of LTO\_N, indicating the presence of Ti–O–N species on the surface (see also Fig. 2b). LTO\_N exhibited not only a sharp peak at 405 eV but also broad split peaks around 400 eV, which can be attributed to Ti–N bonds (Fig. 3b).<sup>63</sup> The XAS analysis indicated that the spatial distribution of N species in LTO\_N is through surface segregation of interstitial N (Ti–O–N) and bulk-substituted N (Ti–N), as described above. More importantly, LTO\_N-HT shows a significant decrease in the peak intensity at 405 eV (Ti–O–N bond) in both the TEY and PFY modes. In contrast, the peak corresponding to the substituted N (Ti–N bond) at 400 eV is well maintained, as shown in Fig. 3b, even after heat treatment. This XAS result validates the peak separation and assignment of the  $\text{N}_2$  profile in TPD (Fig. 1e). Specifically, the low- and high-temperature regions of the  $\text{N}_2$  desorption peak below and above 1100 °C are attributed to interstitial and substitutional N species, respectively. In addition, the XANES shapes of all the LTO samples differed from that of the reference TiN. This indicates that TiN crystals were

not generated inside the LTO samples, and Ti–N bonds were formed by the local substitution of N for O. In summary, the combined study of XPS and XAS elucidated the assignment of  $\text{N}_2$  desorption in the TPD profiles. This underscores the efficacy of TPD in distinguishing between interstitial and substitutional N species within bulk perovskite-type oxides. According to the peak area ratios derived from TPD results, interstitial N and substitutional N in LTO\_N were quantified to be 0.37 wt% and 0.30 wt%, respectively.

### Thermal desorption mechanism of TPD

Finally, the thermal desorption behaviour of the N species from LTO\_TEA and LTO\_N was analysed by *in situ* DRIFTS, as shown in Fig. 4a–d, respectively. LTO\_TEA displayed absorption bands at 3800–2600  $\text{cm}^{-1}$ ,<sup>64</sup> corresponding to OH bonds, and around 2922 and 2853  $\text{cm}^{-1}$ , corresponding to CH bonds (Fig. 4a).<sup>65,66</sup> In contrast, LTO\_N exhibited significantly smaller bands of OH and CH bonds, which is consistent with the observations concluding that organic-derived N species are present in LTO\_TEA but not in LTO\_N. These bands decrease with increasing temperature and become negligible around 600 °C, which is nearly consistent with the desorption temperatures of  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ , NO, and  $\text{N}_2$  (see Fig. 1d and S6a†). Interestingly, an absorption band around 2175  $\text{cm}^{-1}$ , corresponding to the vibration of CO adsorbed on the surface,<sup>67,68</sup> emerged during the heat treatment. Moreover, the peak intensity decreased with increasing temperature after reaching a maximum intensity at approximately 600 °C. The absorption behaviour could be related to CO desorption because the second CO desorption peak starts around 600 °C in Fig. S6a.† Similar IR absorption behaviour and CO desorption in the TPD were also observed for LTO\_N (Fig. 4d and S6b†). A small absorption band at 2000  $\text{cm}^{-1}$ , corresponding to the NO bond,<sup>69–71</sup> was observed for LTO\_TEA (Fig. 4b and S11†). During the heat treatment, the NO vibration band initially increased up to 300 °C and then

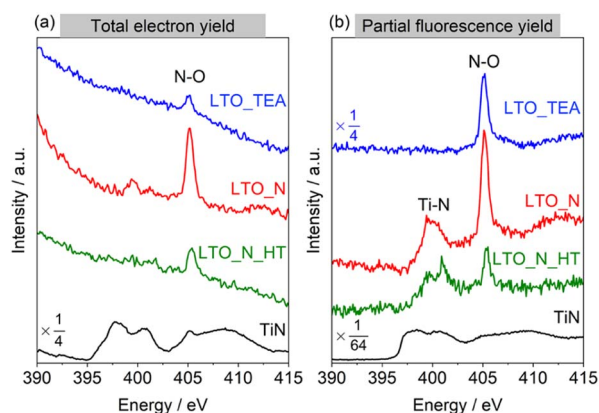


Fig. 3 N K-edge XANES spectra of LTO\_TEA, LTO\_N, LTO\_N-HT, and reference TiN recorded in the (a) TEY and (b) PFY modes.

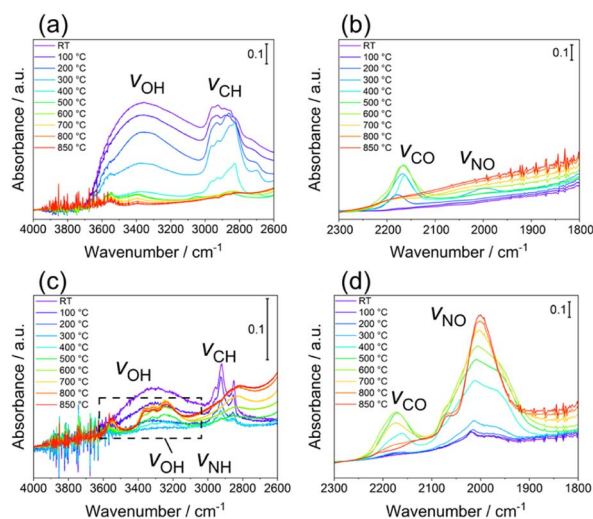


Fig. 4 *In situ* DRIFTS spectra of LTO\_TEA in (a) 4000–2600  $\text{cm}^{-1}$  and (b) 2300–1800  $\text{cm}^{-1}$  regions and LTO\_N in (c) 4000–2600  $\text{cm}^{-1}$  and (d) 2300–1800  $\text{cm}^{-1}$  regions.



decreased with increasing temperature. This absorption behaviour agrees with the TPD profile of LTO\_TEA (see Fig. 1d), because the desorption of NO and N<sub>2</sub> gases began at approximately 300 °C. In the case of LTO\_N (Fig. 4c and d), distinct split bands at approximately 3300 cm<sup>-1</sup>, which are considered to be NH or OH bonds,<sup>65,72,73</sup> and strong absorption at 2000 cm<sup>-1</sup> corresponding to the NO bond were observed. The NH or OH vibration appeared at 400 °C, with intensity increasing with increasing temperature up to 800 °C. The absorption intensity of the NO vibration was also enhanced by the heat treatment. These increases in absorption appear to be related to the desorption behaviour of N<sub>2</sub> in LTO\_N (Fig. 1e). First, the presence of NO bonds in Fig. 4d is consistent with the N<sub>2</sub> desorption in the lower temperature range starting at approximately 600 °C (Fit Peak 1 in Fig. 1e), which is attributed to the Ti–O–N species. Because DRIFTS is a surface-sensitive method,<sup>74</sup> the growth of the NO vibrations may be related to N migration to the top surface of the sample. The N species bound to the oxide in LTO\_N were not desorbed as NO, unlike those in LTO\_TEA. Thus, the Ti–O–N species can probably diffuse to the surface at temperatures below 600 °C but cannot desorb (corresponding to increased IR absorption intensity), and only above 600 °C do they desorb as N<sub>2</sub> gas, as detected by TPD.

## Conclusions

We propose a TPD technique to identify doped N species in perovskite-type oxides. The N 1s XPS spectra of the two types of perovskite-type oxides (LTO\_TEA and LTO\_N) were similar, although they exhibited different physical and chemical properties. In contrast, TPD successfully distinguished the N species in these oxides based on the desorbed gas species and their desorption temperatures, providing accurate quantification in the bulk. The combined study of TPD results with XPS and XANES analyses revealed that it is possible to distinguish between interstitial and substituted N species using the desorption temperature. The distinction is important for perovskite-type oxide applications. Furthermore, *in situ* DRIFTS analysis elucidated the desorption mechanism of the N species. This study is the first to demonstrate the usefulness of the TPD technique for analysing doped N species in oxides. TPD, a bulk analysis, complements the surface-sensitive XPS technique; the combination of TPD and XPS will enable discriminative analysis of surface and bulk structures, providing a powerful method for the further development of functional doped oxide materials. Moreover, TPD technology is not limited to the analysis of nitrogen species in perovskite-type oxides, but is expected to be expanded to other dopants such as hydrogen, sulphur, and phosphorus in various inorganic materials, including oxides, carbides, nitrides, and sulphides. Since the high-temperature vacuum TPD system is a lab-scale setup, it will hopefully be commercially available and widely used in the future.

## Data availability

The data that support the findings of this study are available from the corresponding author upon reasonable request.

## Author contributions

Conceptualization: TY, SY, and HN; methodology: SS, TY, GN, EK, and HN; investigation: SS, GN, and JW; supervision: TY, SY, and HN; writing—original draft: SS; writing—review and editing: TY and HN.

## Conflicts of interest

The authors declare that they have no competing financial interests or personal relationships that may have influenced the work reported in this study.

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