

Cite this: *Chem. Sci.*, 2024, **15**, 9599

All publication charges for this article have been paid for by the Royal Society of Chemistry

Mechanistic investigations of the Fe(II) mediated synthesis of squaraines[†]

Yu Liu, [‡]^a Nathan T. Coles, [‡]^{*}^a Nathalia Cajiao, ^{bc} Laurence J. Taylor, ^a E. Stephen Davies, ^a Alistair Barbour, ^a Patrick J. Morgan, ^a Kevin Butler, ^a Ben Pointer-Gleadhill, ^a Stephen P. Argent, ^a Jonathan McMaster, ^a Michael L. Neidig, ^b David Robinson ^d and Deborah L. Kays ^{*}^a

The scission and homologation of CO is a fundamental process in the Fischer–Tropsch reaction. However, given the heterogeneous nature of the catalyst and forcing reaction conditions, it is difficult to determine the intermediates of this reaction. Here we report detailed mechanistic insight into the scission/homologation of CO by two-coordinate iron terphenyl complexes. Mechanistic investigations, conducted using *in situ* monitoring and reaction sampling techniques (IR, NMR, EPR and Mössbauer spectroscopy) and structural characterisation of isolable species, identify a number of proposed intermediates. Crystallographic and IR spectroscopic data reveal a series of migratory insertion reactions from 1^{Mes} to 4^{Mes} . Further studies past the formation of 4^{Mes} suggest that ketene complexes are formed en route to squaraine 2^{Mes} and iron carboxylate 3^{Mes} , with a number of ketene containing structures being isolated, in addition to the formation of unbound, protonated ketene (8). The synthetic and mechanistic studies are supported by DFT calculations.

Received 23rd February 2024
Accepted 4th May 2024

DOI: 10.1039/d4sc01286k
rsc.li/chemical-science

1. Introduction

The reduction and homologation of CO to obtain high-value organic compounds is a long-standing area of research interest. In industry, the Fischer–Tropsch process utilises heterogeneous catalysts and forcing conditions to convert CO, in the presence of H₂, to hydrocarbon fuels.^{1,2} However, such reactions are unselective, affording products with a range of chain lengths.^{3–5} As such, the reaction of CO with homogeneous species to selectively generate specific organic products remains a tantalising prospect. Over the years, many examples of such reactions have been published from across the periodic table, utilising s-block,^{6–11} p-block,^{12–16} d-block,^{17–25} and f-block elements.^{26–31} Not only do these reactions provide interesting organic products, but the mechanisms of these reactions are also of significant interest. Often, a complex series of steps facilitate this reduction and homologation process.^{32,33}

Understanding these reactions allows us greater insight into not only the chemistry of the elements involved, but also the utilisation of CO on an industrial scale. The Fischer–Tropsch process, being a heterogeneous reaction, is challenging to study mechanistically. The investigation of CO reduction by soluble transition metal complexes can, therefore, provide valuable information about the reactions that may occur in such a system.

The first example of insertion of CO into a low-coordinate Fe(II) *m*-terphenyl complex was reported by Ni and Power.³⁴ Upon exposure of the complex Fe(C₆H₃-2,6-Dipp)₂ (Dipp = 2,6-iPr₂C₆H₃) to an atmosphere of CO, the complex Fe(CO)₂[C(O)C₆H₃-2,6-Dipp]₂ (Dipp = 2,6-iPr₂C₆H₃) was obtained. Subsequently, we reported the selective reduction and homologation of CO by the less bulky Fe(II) *m*-terphenyl complexes (2,6-Ar₂C₆H₃)₂Fe (Ar = Mes (2,4,6-Me₃C₆H₂), 1^{Mes} ; Ar = Xyl (2,6-Me₂C₆H₃), 1^{Xyl}).³⁵ The reaction proceeded at room temperature and 1 atm CO in toluene over 10 days. The final isolated products were a highly unusual squaraine (2^{Mes} , 2^{Xyl} , Scheme 1) featuring broken conjugation between the C₄ and aryl rings, as well as Fe(CO)₅ and an iron carboxylate (3^{Mes} , 3^{Xyl} , Scheme 1). This reaction was particularly noteworthy for the complete scission of the strong C≡O bond (1072 kJ mol^{−1}), which is unusual for a process occurring under such mild conditions. Reactions with the 1-naphthyl (Naph) substituted analogue (2,6-Naph₂C₆H₃)₂Fe (1^{Naph}) afforded an isolable iron carbene complex (CO)₃Fe[C(2,6-Ar₂C₆H₃)OC(O)(2,6-Ar₂C₆H₃)] (Ar = 1-Naph 4^{Naph}), but this species did not show further reactivity to 2

^aSchool of Chemistry, University of Nottingham, University Park, Nottingham, NG7 2RD, UK. E-mail: Deborah.Kays@nottingham.ac.uk; Nathan.Coles@nottingham.ac.uk

^bInorganic Chemistry Laboratory, Department of Chemistry, University of Oxford, South Parks Road, Oxford, OX1 3QR, UK

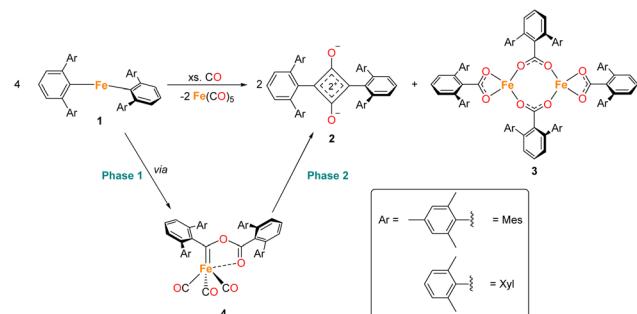
^cDepartment of Chemistry, University of Rochester, Rochester, New York 14627, USA

^dDepartment of Chemistry and Forensics, School of Science and Technology, Nottingham Trent University, Nottingham, NG11 8NS, UK

[†] Electronic supplementary information (ESI) available. CCDC 2334282–2334296.

For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d4sc01286k>

[‡] Authors contributed equally.



Scheme 1 General reaction scheme for the reaction between **1** in the presence of CO, where Ar = Mes or Xyl. The two phases of the reaction are shown, phase 1 covers the reaction of **1** to **4**, phase 2 covers the reaction from **4** to **2**, **3** and Fe(CO)₅.

and **3**. Regardless, analogous iron carbene complexes (**4^{Mes}**, **4^{Xyl}**) were proposed as intermediates in the reaction between **1^{Mes}** and **1^{Xyl}** with CO.

Herein, we report detailed mechanistic investigations of the homologation and scission of CO by a two-coordinate Fe(II) complex using *in situ* monitoring techniques (IR, NMR and EPR spectroscopy), frozen-solution Mössbauer (MB) spectroscopy and structural characterisation of intermediates *via* single crystal X-ray crystallography (scXRD). These experimental results are supported by DFT studies. From this, we propose a plausible mechanistic pathway, which highlights the unusual reactivity that can be facilitated using sterically demanding ligands which enforce low coordination environments at a metal centre.

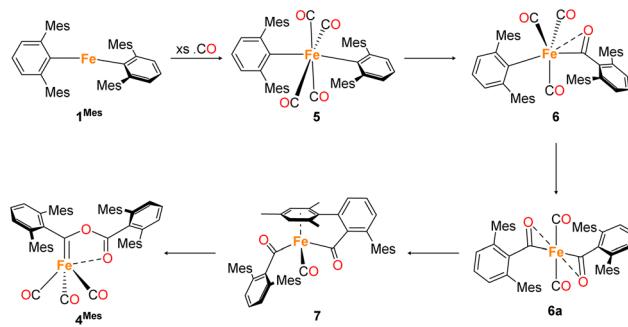
2. Results and discussion

2.1 Overview

During the course of reaction monitoring (see discussion below) it became clear that the transformation of **1^{Mes}**–**3^{Mes}**, and Fe(CO)₅ proceeds in two distinct stages (Scheme 1). In phase 1, **1^{Mes}** reacts with CO, *via* several reactive intermediates, to afford an iron–carbene complex **4^{Mes}** (Scheme 1), which becomes the dominant species in solution. Phase 2 of the reaction involves **4^{Mes}** reacting with further equivalents of CO resulting in the formation of a number of species, and eventually the final products **2^{Mes}**, **3^{Mes}** and Fe(CO)₅. For simplicity, we will discuss these two phases of the reaction separately. Sections 2.2 and 2.3 will focus on phase 1, covering all intermediates and observations up to the formation of **4^{Mes}**. Sections 2.4–2.7 will cover phase 2, looking at potential intermediates and pathways between **4^{Mes}** and the final products.

2.2 Spectroscopic analysis of phase 1

In situ IR spectroscopy was conducted to monitor the transformation of the iron complex Fe(C₆H₃-2,6-Mes₂)₂ (**1^{Mes}**) in the presence of CO over the course of the reaction. We postulate that the first step is the coordination of four CO molecules to **1^{Mes}** affording Fe(CO)₄(C₆H₃-2,6-Mes₂)₂ (**5**, Scheme 2). IR spectra of the reaction between **1^{Mes}** and *ca.* 1 atm of CO in toluene were recorded at two-minute intervals. Within the first



Scheme 2 Proposed reaction pathway for formation of **4^{Mes}** from **1^{Mes}**. Note that **6a** has only been observed spectroscopically in solution, all other species have been isolated and characterised by single crystal X-ray diffraction (See Section 2.3).

hour of the reaction, several carbonyl-containing species form which are subsequently consumed (Fig. 1). Initial intense signals at $\nu(\text{CO}) = 2014 \text{ cm}^{-1}$, 2004 cm^{-1} and 1979 cm^{-1} were accompanied by less intense signals at $\nu(\text{CO}) = 2076 \text{ cm}^{-1}$, 1942 cm^{-1} and 1935 cm^{-1} . Following this, less intense peaks are observed at $\nu(\text{CO}) = 2063 \text{ cm}^{-1}$ and 2035 cm^{-1} after approximately 30 minutes. The signals at $\nu(\text{CO}) = 2076 \text{ cm}^{-1}$, 2014 cm^{-1} and 1979 cm^{-1} can be attributed to **6** which has been isolated and characterised independently in the solid state (See Section 2.3). After *ca.* 40 minutes of *in situ* monitoring the strongest IR bands occur at $\nu(\text{CO}) = 2049 \text{ cm}^{-1}$, 1978 cm^{-1} and 1965 cm^{-1} , which are assigned to **4^{Mes}**. These are similar to the previously reported **4^{Naph}** [$\nu(\text{CO}) = 2043 \text{ cm}^{-1}$, 1972 cm^{-1} and 1954 cm^{-1}], a structural analogue of **4^{Mes}**.³⁵ It has not been possible to assign the remaining peaks to specific species, but they are postulated to be metal carbonyl complexes.

The initial stages of the reaction were also monitored *via* ¹H NMR spectroscopy. **1^{Mes}** is a paramagnetic complex, and displays resonances over the range +80 to –180 ppm. However, upon introduction of CO, several new ¹H NMR signals appear in the range 0–8 ppm. This is attributed to the formation of diamagnetic 18e[–] iron complexes. Forty minutes after addition of CO, several species form, resulting in a set of overlapping signals that cannot be resolved from one another (Fig. 2). As the reaction proceeds four signals become dominant at $\delta_{\text{H}} = 2.31 \text{ ppm}$, 2.23 ppm , 1.83 ppm and 1.65 ppm , consistent with a species featuring terphenyl moieties in two separate environments. The IR and ¹H NMR spectroscopic data indicate one major species is formed, **4^{Mes}** (see Section 2.3). Full consumption of **1^{Mes}** takes *ca.* 36 h, with concomitant formation of a paramagnetic species (See Fig. S45–S47, ESI[†]).

The use of ¹³CO allows the observation of intermediates by ¹³C{¹H} NMR spectroscopy, showing the simultaneous formation of several compounds which contain CO in the form of acyl/carbene functional groups (250–270 ppm), metal-bound CO (220–200 ppm) and esters (176 ppm). Ni and Power have previously reported Fe(CO)₂[C(O)C₆H₃-2,6-Dipp]₂ (Dipp = 2,6-iPr₂C₆H₃) which showed carbonyl and acyl signals at $\delta_{\text{C}} = 214.8 \text{ ppm}$ and 258.5 ppm , respectively. In our work, integration of the ¹³C{¹H} NMR of spectra after *~*30 minutes of being placed under an atmosphere of ¹³CO shows that the signals at $\delta_{\text{C}} =$



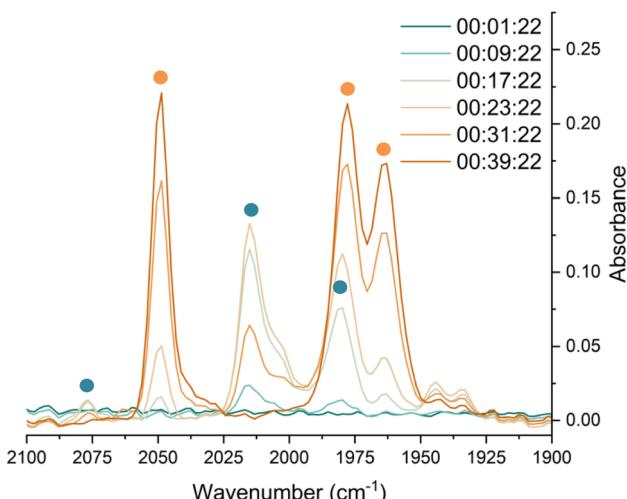


Fig. 1 IR spectra for the first 40 minutes of the reaction between 1^{Mes} and CO in toluene. CO added to the reaction after six minutes. The colour gradient goes from dark turquoise (earliest) to dark orange (latest). Orange circles highlight the signals for 4^{Mes} , turquoise circles highlight the signals attributed to 6.

259.5 ppm and 206.0 ppm occur in a 1 : 3 ratio, while the signals at $\delta_{\text{C}} = 257.6$ ppm and 214.7 ppm occur in a 1 : 1 ratio (Fig. S48, ESI ‡). This suggests the formation of structures of the type 6 and 6a, respectively, en route to 4^{Mes} .

We note that, when monitoring the reaction using a ReactIR spectrometer, conversion to the carbene (4^{Mes}) is complete within 40 minutes, but this takes *ca.* 36 hours when monitoring by *in situ* ^1H NMR spectroscopy. This is presumably due to a smaller headspace of CO, smaller interfacial surface area and, less efficient mixing. Paramagnetic species are also formed after 4^{Mes} , which hinders further *in situ* monitoring by NMR spectroscopy.

2.3 Synthesis and structural characterisation of iron complexes observed during phase 1

Through careful control of the reaction conditions *via* solvent choice and reaction monitoring, it is possible to isolate the

iron–carbene $(\text{CO})_3\text{Fe}[\text{C}(2,6\text{-Mes}_2\text{C}_6\text{H}_3)\text{OC}(\text{O})(2,6\text{-Mes}_2\text{C}_6\text{H}_3)]^{\text{4}^{\text{Mes}}}$, (Scheme 2), in preparative quantities. The reaction between 1^{Mes} and an atmosphere of CO in hexane affords an orange solution from which 6 precipitates (Scheme 2, see discussion below), after *ca.* 10 minutes. The reaction mixture is then stirred until complete redissolution of 6 occurs, followed by an additional 10 minutes of stirring. Filtration of the dark orange-red solution followed by removal of the volatiles allows the isolation of 4^{Mes} in 95% yield. 4^{Mes} has been characterised by ^1H and $^{13}\text{C}[^1\text{H}]$ NMR spectroscopy (Fig. S1 and S2, ESI ‡), IR spectroscopy, mass spectrometry and elemental analysis. The NMR spectra confirm that the signals observed at $\delta_{\text{H}} = 2.31$ ppm, 2.23 ppm, 1.83 ppm and 1.65 ppm in the ^1H NMR spectroscopic monitoring of the reaction between CO and 1^{Mes} (Fig. 2) are due to 4^{Mes} . The ATR-FTIR spectrum of 4^{Mes} (Fig. S18 and S19, ESI ‡) displays three strong stretches which are observed at $\nu(\text{CO}) = 2046\text{ cm}^{-1}$, 1974 cm^{-1} and 1959 cm^{-1} corresponding to the metal-bound CO groups, with a less intense stretch at $\nu(\text{CO}) = 1612\text{ cm}^{-1}$ for the carboxyl group bound to the carbene. Again, this is consistent with 4^{Mes} being the major species present after 40 minutes during *in situ* IR reaction monitoring (Fig. 1, Section 2.2).

4^{Mes} has also been characterised by single crystal X-ray diffraction. Two solvatomorphs have been isolated; 4^{Mes} (Fig. 3) and $4^{\text{Mes}}\cdot\text{Et}_2\text{O}$ (see ESI, Fig. S36 and Table S2 ‡), both of which have been grown from concentrated Et_2O solutions at low temperature. The Fe1–C1 bond length 4^{Mes} is near identical to the analogous distance for the previously published 4^{Naph} [1.8395(14) Å and 1.840(3) Å, respectively] suggesting the presence of an $\text{Fe}=\text{C}$ bond.³⁵

Whilst intermediate species between 1^{Mes} and 4^{Mes} are highly reactive, it has been possible to crystallise proposed intermediates of this transformation (Scheme 2). The reaction between 1^{Mes} and 1 atm of CO in benzene resulted in a colour change from yellow to orange. Five minutes after the change in colour, the reaction mixture was flash frozen in liquid N_2 . The solvent was then sublimed off under vacuum, resulting in the formation of a red solid. This solid was then extracted three times with iso-hexane, keeping the extractions separate and

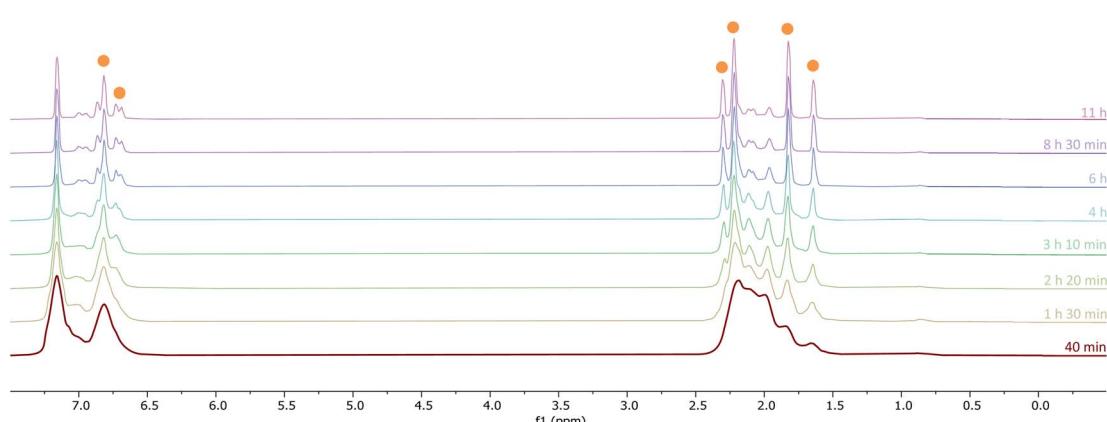


Fig. 2 ^1H NMR spectra (–1 to 8 ppm) of the reaction between CO and 1^{Mes} (C_6D_6 , 1 atm CO, room temperature) recorded between 40 minutes and 11 hours of reaction. Orange circles highlight the signals for 4^{Mes} .



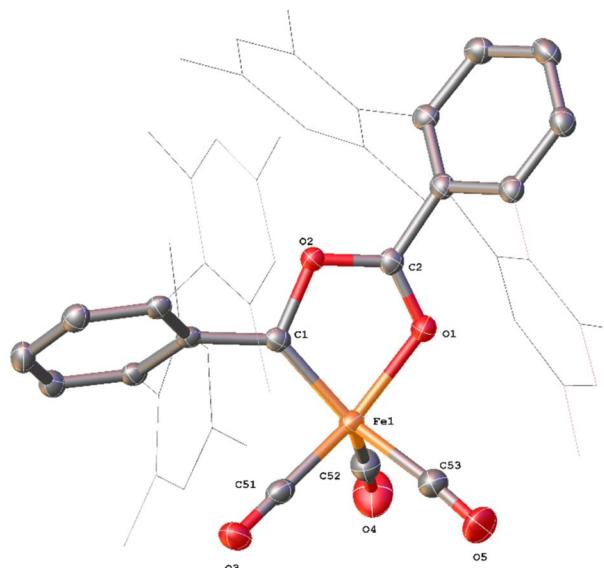


Fig. 3 View of the metal complex from the crystal structure of 4^{Mes} with anisotropic displacement ellipsoids set at 50% probability. Mesityl groups shown as wireframe, hydrogen atoms, two co-crystallised diethyl ether and a second equivalent of 4^{Mes} have been omitted for clarity. Selected bond distances (Å) and angles (°) for 4^{Mes} shown: Fe1–C1 1.8395(14), Fe1–O1 1.9470(11), Fe1–C51 1.7549(16), Fe1–C52 1.828(2), Fe1–C53 1.8382(19), Fe1–C1–O1 80.81(5), O1–C2–O2 118.77(13).

these solutions were cooled to 8 °C for 48 hours. From these solutions, crystals of $\text{Fe}(\text{CO})_4(\text{C}_6\text{H}_3\text{-}2,6\text{-Mes}_2)_2$ (**5**), $\text{Fe}(\text{CO})_3[\text{C}(\text{O})\text{C}_6\text{H}_3\text{-}2,6\text{-Mes}_2](\text{C}_6\text{H}_3\text{-}2,6\text{-Mes}_2)$ (**6**) and $\text{Fe}(\text{CO})[\text{C}(\text{O})\text{C}_6\text{H}_3\text{-}2,6\text{-Mes}_2]_2$ (**7**) suitable for X-ray crystallography were obtained (Fig. 4–6). These represent the coordination of four CO molecules to 1^{Mes} , followed by sequential migratory insertion reactions. Crystals of these complexes were isolated from mixtures containing several metal-containing species, therefore, it has not been possible to isolate **5**, **6** or **7** in sufficient purity or quantities for full analysis.

The solid-state structure of **5** (Fig. 4), features a rare *trans*-arrangement of the terphenyl ligands,^{36–40} which is closer to linearity than 1^{Mes} [C1–Fe1–C25 angle of 174.57(10)°].⁴¹ Complex **5** shows a near-octahedral geometry at the iron(II) centre with bent CO ligands due to the steric hindrance of the terphenyl groups.

Complex **6** (Fig. 5) features one *m*-terphenyl ligand, three carbonyls and an acyl ligand formed through the migratory insertion of one CO ligand into the Fe–C bond of the second *m*-terphenyl ligand. The acyl ligand coordinates in an η^2 -bonding mode, affording an 18e[–] complex. These bond lengths and angles are similar to those observed for the diacyl complex $\text{Fe}(\text{CO})_2[\text{C}(\text{O})\text{C}_6\text{H}_3\text{-}2,6\text{-Dipp}_2]_2$ which features a longer Fe–C bond than in **6** (1.8964(18) Å vs. 1.872(3) Å) but a decreased Fe–O distance (2.0229(14) Å vs. 2.110(3) Å) for the acyl group.³⁴ Reaction of 1^{Mes} (200 mg) in iso-hexane (20 mL) under an atmosphere of CO yielded an orange precipitate (15% yield) corresponding to intermediate **6**. ATR-FTIR analysis of **6** revealed CO stretches at 2077 cm^{–1}, 2013 cm^{–1} and 1975 cm^{–1}

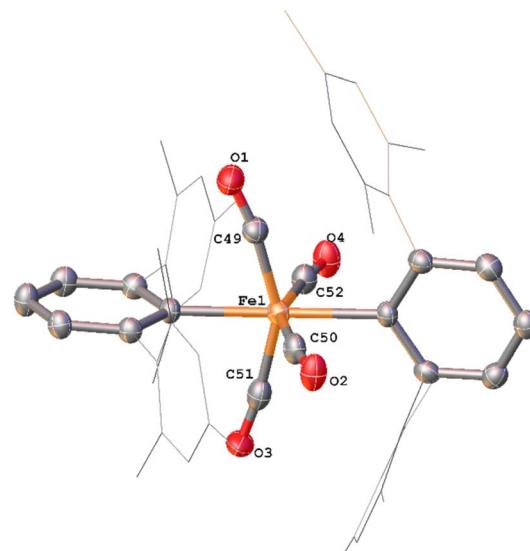


Fig. 4 View of the metal complex from the crystal structure of **5** with anisotropic displacement ellipsoids set at 50% probability. Mesityl groups shown as wireframe and hydrogen atoms have been omitted for clarity. Selected bond distances (Å) and angles (°) for **5** shown: Fe1–C1 2.156(3), Fe1–C25 2.171(3), Fe1–C49 1.808(3), Fe1–C50 1.821(3), Fe1–C51 1.812(4), Fe1–C52 1.817(4), C1–Fe1–C25 174.58(11), C49–Fe1–C51 152.64(15), C50–Fe1–C52 153.63(15).

(Fig. S22 and 23, ESI[‡]) which correspond with those observed during the *in situ* measurements in toluene (Fig. 1; $\nu(\text{CO}) = 2076 \text{ cm}^{-1}$, 2014 cm^{-1} and 1979 cm^{-1}). The carbonyl stretch

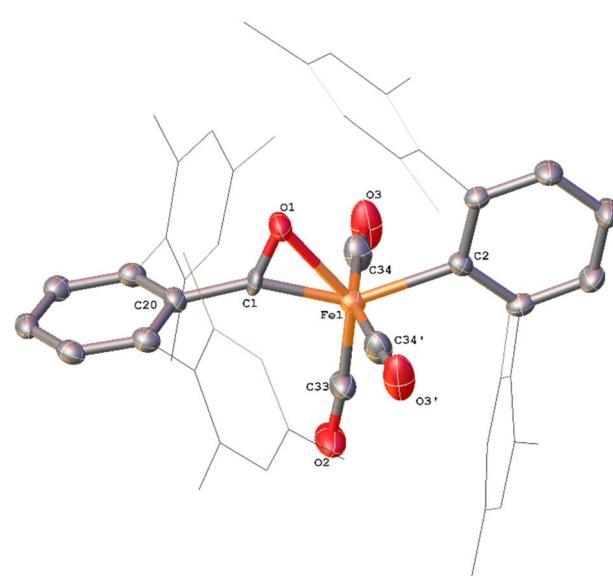


Fig. 5 View of the metal complex from the crystal structure of **6** with anisotropic displacement ellipsoids set at 50% probability. Mesityl groups shown as wireframe and hydrogen atoms have been omitted for clarity. The following symmetry operations were used to generate the marked atoms: $+x, 3/2 -y, +z$. Selected bond distances (Å) and angles (°) for **6** shown: C1–O1 1.221(4), C33–O2 1.148(5), C34–O3 1.130(4), Fe1–C1 1.872(3), Fe1–C2 2.104(4), Fe1–O1 2.110(3), C20–C1–O1 124.6(3), Fe1–C1–O1 83.2(2).

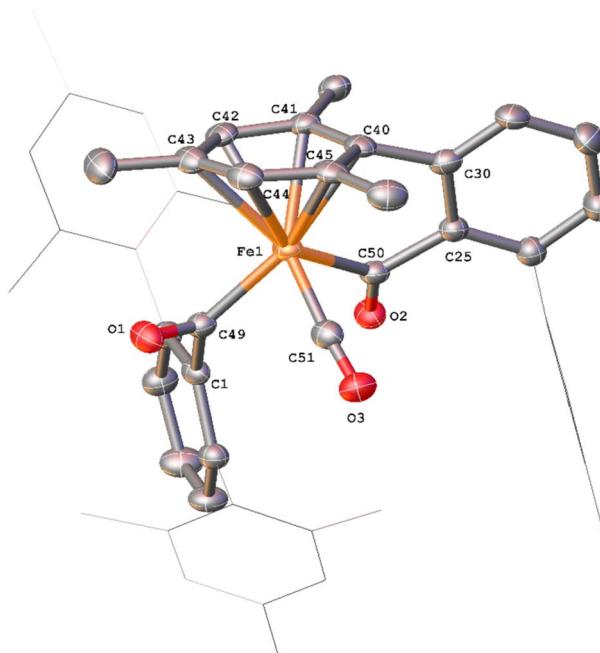


Fig. 6 View of the metal complex from the crystal structure of 7 with anisotropic displacement ellipsoids set at 50% probability. Mesityl groups shown as wireframe and hydrogen atoms have been omitted for clarity. Selected bond distances (Å) and angles (°) for 7 shown: C49–O1 1.221(3), C50–O2 1.204(3), Fe1–C49 1.997(2), Fe1–C50 1.986(2), C1–C49–O1 116.24(19), C25–C50–O2 120.7(2), Fe1–C49–C1 129.13(15), Fe1–C50–C25 111.44(15), Fe1–C49–O1 114.59(17), Fe1–C50–O2 127.87(18).

corresponding to the acyl ligand in this complex is observed at 1615 cm^{-1} . Mass spectrometric analysis of **6** using MALDI-TOF allowed the observation of the $[\text{M} - 2(\text{CO})]^+$ ion (see ESI†). When dissolving **6** to obtain an NMR spectrum, **4^{Mes}** was detected within 10 minutes, even though there was no **4^{Mes}** present in the ATR-IR spectrum of the solid, demonstrating their intrinsic high reactivity. *In situ* $^{13}\text{C}\{^1\text{H}\}$ NMR studies of the reaction between **1^{Mes}** and ^{13}CO signals give signals at $\delta_{\text{C}} = 259.5\text{ ppm}$ and 206.0 ppm (integral ratio of 1 : 3, Fig. S48, ESI†), which have been tentatively assigned to **6**.

The solid state structure of **7** (Fig. 6) features one CO and two η^1 -acyl ligands, one of which also binds to the Fe *via* an η^6 -mesityl group, affording an 18e^- metal centre. This differs substantially to the Dipp-substituted species $\text{Fe}(\text{CO})_2(\text{C}(\text{O})\text{C}_6\text{H}_3\text{-2,6-Dipp}_2)_2$, in which the complex bears two η^2 -acyl moieties and two CO ligands.³⁴ This difference is attributed to the lower steric demands of the mesityl substituents in **7** relative to the bulky 2,6-diisopropylphenyl moieties. As a result, the acyl ligands of **7** show significantly different bond lengths and angles relative to **6**, consistent with η^1 -coordination, most notably increased Fe–C(acyl) bond lengths [1.997(2) Å (Fe1–C49) and 1.986(2) Å (Fe1–C50)]. We propose that **7** is likely only observed in the solid state as, unlike **6**, there are no signals in the ^{13}CO -labelled $^{13}\text{C}\{^1\text{H}\}$ NMR spectrum that can be assigned to **7**. In a solution saturated with CO we propose **7** exists as **6a** (Scheme 2) with two bound CO ligands, akin to $\text{Fe}(\text{CO})_2(\text{C}(\text{O})\text{C}_6\text{H}_3\text{-2,6-Dipp}_2)_2$. This is based on similarities within the ^{13}CO -labelled $^{13}\text{C}\{^1\text{H}\}$ NMR spectra where

signals were observed at $\delta_{\text{C}} = 257.6\text{ ppm}$ and 214.7 ppm (1 : 1 ratio) for **6a**, *cf.* acyl: $\delta_{\text{C}} = 258.5\text{ ppm}$, carbonyl: 214.8 ppm for $\text{Fe}(\text{CO})_2(\text{C}(\text{O})\text{C}_6\text{H}_3\text{-2,6-Dipp}_2)_2$.³⁴

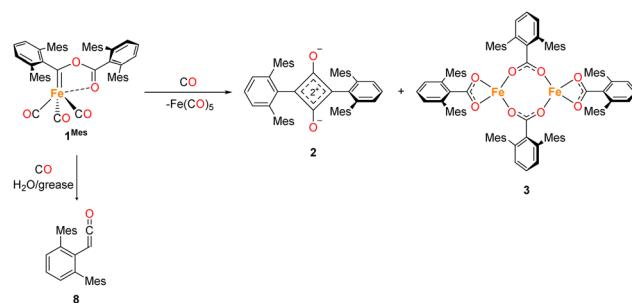
To further probe these structural insights, DFT calculations were employed. The calculations show a low barrier to the formation of **6** from **5** (21.9 kJ mol⁻¹), while a further migratory insertion reaction occurs to form **6a** (Fig. S86, ESI†) before rapid formation of **7**, with a barrier height relative to **5** of only 24.1 kJ mol⁻¹. The reaction then proceeds further to, followed by rapid reaction to further intermediates. DFT calculations suggest the barrier to formation of **6a** from **6** is *ca.* 70 kJ mol⁻¹ without solvent, confirming the relative stability of **6**.

2.4 Spectroscopic analysis of phase 2

Reaction monitoring of phase 2 towards the final products (Scheme 3) squaraine (**2^{Mes}**), $\text{Fe}(\text{CO})_5$ and $\text{Fe}_2[\text{O}_2\text{C}(2,6\text{-MesC}_6\text{H}_3)]_4$ (**3**), which takes up to an additional 8 days post formation of **4^{Mes}**, has proven more challenging. Through ReactIR, it was impossible to avoid ingress of small amounts of water and/or oxygen into the flask, even when a continuous positive atmosphere of CO was employed, leading to side reactions of the highly reactive intermediates. Additionally, **3** is an insoluble, paramagnetic solid that precipitates during the course of the reaction, which interferes with *in situ* NMR spectroscopic monitoring.

To obtain an endpoint IR spectrum upon completion of the reaction, **1^{Mes}** was reacted in toluene under *ca.* 1 atm of CO for 7 days. The reaction was filtered to remove insoluble **3** and an IR spectrum of the resulting solution was obtained (Fig. S45, ESI†). The observed peaks were assigned as **2** (1674 cm^{-1}), $\text{Fe}(\text{CO})_5$ (1996 cm^{-1} , 2024 cm^{-1}) and, at 2097 cm^{-1} , a ketene ($\text{O}=\text{C}=\text{C}(\text{H})\text{C}_6\text{H}_3\text{-2,6-Mes}_2$, **8**) which we propose results from reaction with moisture and/or silicone grease. Ketene **8** has been characterised in reaction mixtures by ^1H NMR, $^{13}\text{C}\{^1\text{H}\}$ NMR and IR spectroscopy, and mass spectrometry (see ESI, Fig. S6, S7, S31, S32 and S50†), although it has not been possible to isolate this as the sole product.

To monitor phase 2, by IR spectroscopy, a toluene solution of **4^{Mes}** was reacted with CO in a J. Young reaction flask. Periodically, an aliquot of the reaction was filtered to remove insoluble **3** and transferred to a sealable IR cell. Over the course of 8 days, signals corresponding to **4^{Mes}** ($\nu(\text{CO}) = 2049\text{ cm}^{-1}$, 1978 cm^{-1})



Scheme 3 General reaction scheme post the formation of **4^{Mes}** and the observed degradation product upon reaction with moisture/grease.



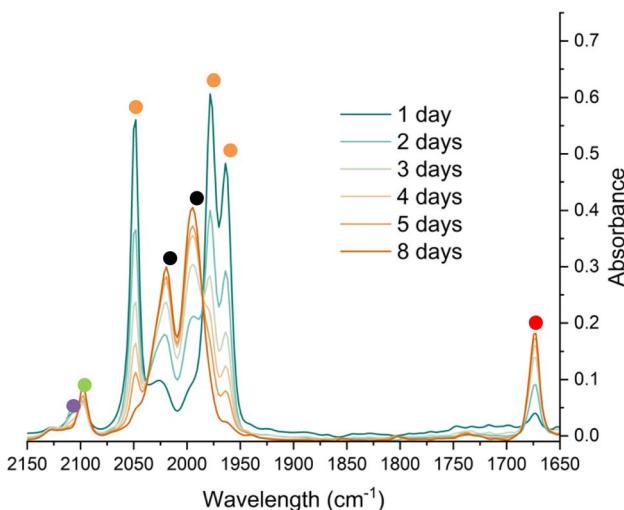


Fig. 7 IR spectroscopic monitoring using a Harrick cell of the reaction of 4^{Mes} under an atmosphere of CO in toluene over 8 days. Orange circles highlight the signals for 4^{Mes} , red circles highlight the signals for 2^{Mes} . Black circles highlight the signals for $\text{Fe}(\text{CO})_5$. The green circle highlights the signal for 8 . The purple circle highlights a signal attributed to a proposed Fe -ketenyl complex, 9 (see Fig. 8 for proposed structure). See ESI[‡] for a zoom in of the region between 2150 – 2075 cm^{-1} .

and 1965 cm^{-1}) are consumed (Fig. 7) and are replaced by signals for 2 ($\nu(\text{CO}) = 1674$ cm^{-1}) and ketene 8 ($\nu(\text{CO}) = 2097$ cm^{-1}). After 24 hours, an additional small signal is observed at 2107 cm^{-1} (Fig. S51, ESI[‡]), proposed to be a ketenyl-iron complex. This signal reaches a maximum intensity after *ca.* 48 hours, then decays. Signals relating to $\text{Fe}(\text{CO})_5$ and 2^{Mes} are observed from the first spectra obtained after 24 hours.

^1H NMR spectroscopy was also used to investigate phase 2 of the reaction. As previously mentioned, 3^{Mes} precipitates during the reaction, hindering NMR measurements. Additionally, *in situ* NMR spectroscopic monitoring in a J. Young NMR tube

leads to a different distribution of products after 4^{Mes} . This is attributed to the small headspace (see Section 2.2) leading to lower CO concentrations and side reactions. To circumvent this, aliquots of a larger reaction were filtered from a J. Young reaction flask into NMR tubes at different time points and the ^1H NMR spectrum collected (Fig. 8). Signals for 2^{Mes} are observed at $\delta_{\text{H}} = 1.98$ ppm and 2.20 ppm within 24 hours of the start of the reaction. 8 was also observed ($\delta_{\text{H}} = 2.07$ ppm , 2.22 ppm), which is consistent with IR spectroscopic measurements. Low intensity signals at $\delta_{\text{H}} = 2.36$ ppm , 2.34 ppm and 2.14 ppm can be seen on day 2 and are consumed by day 6 (purple circles, Fig. 8). This is, again, in line with the IR spectroscopic monitoring where a ketene-type signal reaches a maximum concentration on day 3 and is subsequently consumed. Given the resolution of the peaks in the ^1H NMR spectra, we conclude that this is likely an 18e^- , diamagnetic complex. Complex 9 shown in Fig. 8 is the type of structure we propose that we are observing, however, this is only a tentative assignment based on the limited data and ketene containing complexes isolated (see Section 2.5). Additionally, a second diamagnetic complex, 10 (blue circles, Fig. 8) is consistently formed in small quantities. Full characterisation of 10 will be discussed further in Section 2.5.

As $\text{Fe}(\text{CO})_5$ is a product, the reaction may proceed *via* disproportionation of an $\text{Fe}(\text{i})$ complex. To probe the two potential pathways, EPR and Mössbauer (MB) spectroscopies were employed to gain further insight into any intermediary complexes. During the course of the EPR monitoring we observe two signals at room temperature centred at *ca.* $g_{\text{iso}} = 2.038$ and $g_{\text{iso}} = 2.003$ (Fig. S57, ESI[‡]). The signal at $g_{\text{iso}} = 2.003$ is consistent with the previously synthesised radical anion $[2]^{+}$ (Fig. S58, ESI[‡]).³⁵ The signal at $g_{\text{iso}} = 2.038$ appears as a singlet. Freezing the solution did not result in anisotropic splitting of this signal, indicating it is not due to an $\text{Fe}(\text{i})$ complex (Fig. S59, ESI[‡]).

MB spectroscopy required a higher concentration reaction solution relative to the IR and NMR spectroscopic experiments (8-fold increase, 0.23 mol L^{-1}) to enable freeze-quenched

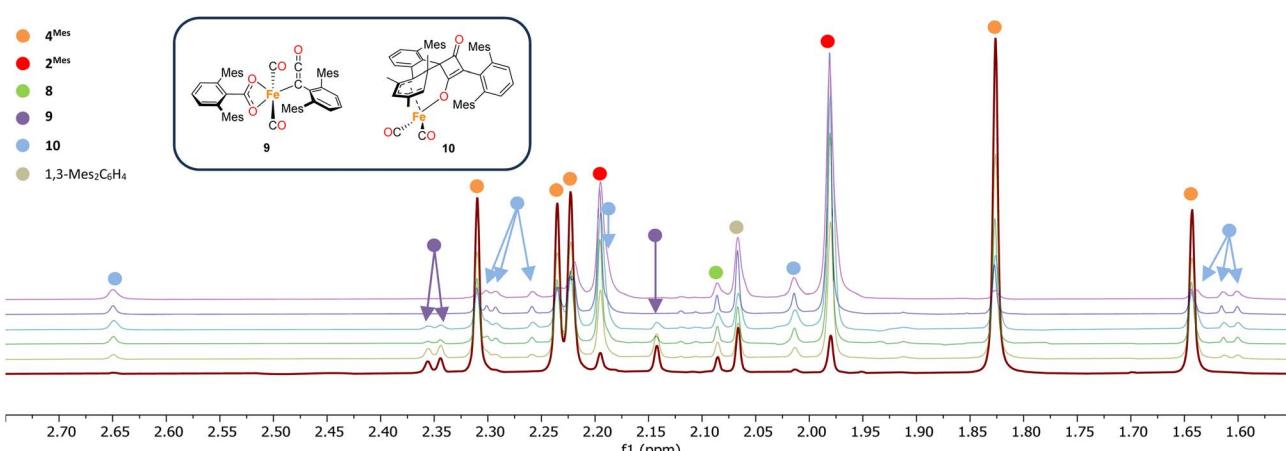


Fig. 8 ^1H NMR spectra of the reaction of 4^{Mes} with CO in C_6H_6 . Spectra collected on after reacting for 1 (red), 2 (yellow), 3 (green), 4 (light blue), 5 (dark blue) and 8 (purple) days. Time increases from bottom to top. Corresponding *para* methyl signals for 8 and $1,3\text{-Mes}_2\text{C}_6\text{H}_4$ overlap with other signals and are not shown. See the ESI for full spectrum (Fig. S56[‡]) and further details on how the data was collected. Inset, structures of proposed Fe-ketenyl intermediate 9 and cyclobuteneone byproduct 10 .



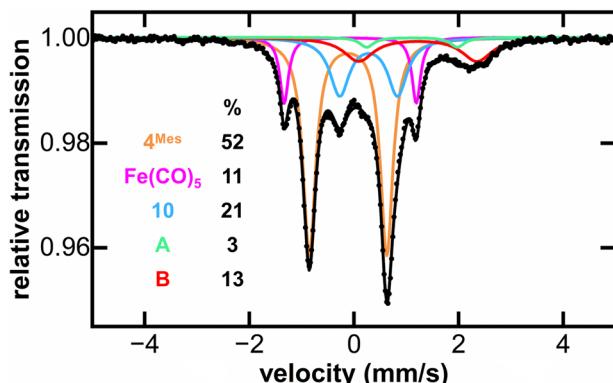


Fig. 9 Freeze-trapped 80 K Mössbauer spectrum of the *in situ* formed iron species upon reaction of **4^{Mes}** with CO at 24 hours of reaction. The individual Mössbauer components are identified as **4^{Mes}** (orange, $\delta = -0.10 \text{ mm s}^{-1}$, $|\Delta E_Q| = 1.56 \text{ mm s}^{-1}$) $\text{Fe}(\text{CO})_5$ (pink, $\delta = -0.08 \text{ mm s}^{-1}$, $|\Delta E_Q| = 2.54 \text{ mm s}^{-1}$), **10** (light blue, $\delta = 0.30 \text{ mm s}^{-1}$, $|\Delta E_Q| = 1.10 \text{ mm s}^{-1}$), **A** (light green, $\delta = 1.10 \text{ mm s}^{-1}$, $|\Delta E_Q| = 1.79 \text{ mm s}^{-1}$) and **B** (red, $\delta = 1.21 \text{ mm s}^{-1}$, $|\Delta E_Q| = 2.33 \text{ mm s}^{-1}$). Raw data are shown as black dots, total fit as a black line, and individual components as colored lines. Further reaction time points at 2, 3, 4 and 5 days are given in the ESI (Fig. S61).[‡]

solution monitoring with natural abundance iron. Freeze-quenched MB samples of the reaction of **4^{Mes}** with CO were collected over the course of the 5 days reaction. Within the first 24 h of reaction, the 80 K MB spectroscopy (Fig. 9) indicated the consumption of nearly 50% of **4^{Mes}** ($\delta = -0.10 \text{ mm s}^{-1}$, $|\Delta E_Q| = 1.56 \text{ mm s}^{-1}$) together with the formation of the $\text{Fe}(\text{CO})_5$ product ($\delta = -0.08 \text{ mm s}^{-1}$, $|\Delta E_Q| = 2.54 \text{ mm s}^{-1}$).⁴² Note that **3** is not observed in these spectra as it is filtered away prior to collection. Over the course of the next 4 days, further generation of $\text{Fe}(\text{CO})_5$ is observed as **4^{Mes}** is consumed (Fig. S59, ESI[‡]). Three additional iron species are also observed (**10**, **A** and **B**) that increase over the course of the reaction and are attributed to decomposition products, one of which is **10** that is observed by ¹H NMR spectroscopy (Fig. 8). Signals for **10**, **A** and **B** are also observed when reacting **4^{Mes}** in the absence of an atmosphere of CO (Fig. S85, ESI[‡]). While the identities of **A** and **B** remain unknown, the MB parameters of **10** are consistent with a diamagnetic, low-spin iron(II) species.^{43,44}

2.5 Synthesis and structural characterisation of iron complexes in phase 2

Complex **10**, which is commonly observed as a minor product, has been isolated as a mixture with **2^{Mes}**. Through *in situ* reaction monitoring, **10** was observed in larger quantities (relative to **2^{Mes}**) when the reaction was performed in an NMR tube. This was proposed to be due to less available CO in the reaction solution. Performing the reaction without an atmosphere of CO, greatly increases the proportion of **10**. Both **10** and **2^{Mes}** precipitate readily from *n*-hexane, both forming intensely orange/red crystals. Therefore, **10** has not been isolated pure for IR and NMR spectroscopic analysis. However, structural characterisation of **10** was possible by scXRD (Fig. 10). **10** contains a dearomatized mesityl ring, which is bound to an $[\text{Fe}(\text{CO})_2]^{2+}$

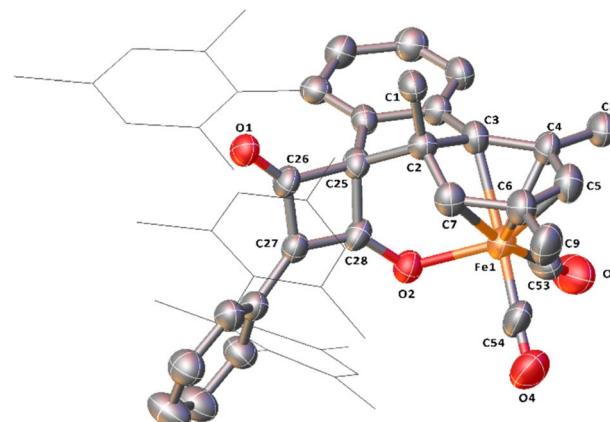


Fig. 10 View of the metal complex from the crystal structure of **10** with anisotropic displacement ellipsoids set at 50% probability. Mesityl groups shown as wireframes and hydrogen atoms have been omitted for clarity. Selected bond distances (Å) for **10** shown: C2–C3 1.535(8), C2–C7 1.496(8), C3–C4 1.399(8), C4–C5 1.414(9), C5–C6 1.404(9), C6–C7 1.394(9), C25–C26 1.548(9), C25–C28 1.538(9), C26–C27 1.450(9), C27–C28 1.375(9), C26–O1 1.213(7), C28–O2 1.270(7), Fe1–C53 1.759(9), Fe1–C54 1.738(9), Fe1–O1 2.011(4).

core as a cyclohexadienyl group with the anionic O of the cyclobutene core also bound to the Fe. The dearomatization is the result of the remaining C of the formerly aromatic cyclohexadienyl ring forming a spirocycle with the cyclobutene core. The cyclohexadienyl bond lengths are within error of other similar complexes,^{45–48} with the delocalised C–C bonds shorter than the single C–C bonds. The metal-bound cyclobutene contains comparable bond lengths to a free cyclobutene synthesised by Heimgartner *et al.*⁴⁹ The structure fits well with the ¹H NMR signals obtained during reaction monitoring. Nine signals are obtained for the methyl groups in **10** ($\delta_{\text{H}} = 0.70 \text{ ppm}$, 1.60 ppm (two overlapping signals), 1.62 ppm , 2.01 ppm , 2.18 ppm , 2.25 ppm , 2.30 ppm and 2.65 ppm) due to steric hindrance preventing free rotation of one of the mesityl rings. The cyclohexadienyl group also has two distinct signals at $\delta_{\text{H}} = 3.54 \text{ ppm}$ and 4.09 ppm for the protons of the cyclohexadienyl group, which is in agreement with other similar complexes.^{48,50,51} The ¹³C{¹H} NMR spectrum for **10** also contains a number of indicative signals between $\delta_{\text{C}} = 50$ – 115 ppm , clearly showing the signals for the cyclohexadienyl moiety, as well as two signals for the cyclobutene at $\delta_{\text{C}} = 89.3 \text{ ppm}$ and 123.2 ppm . Two M–CO signals ($\delta_{\text{C}} = 204.6 \text{ ppm}$ and 207.4 ppm) and a further two distinct C–O signals for the cyclobutene core ($\delta_{\text{C}} = 181.7 \text{ ppm}$ and 191.1 ppm) are also observed. These signals compare well to those reported by Heimgartner *et al.*⁴⁹ Further confirmation for **10** comes from ¹³CO-labelling (see Section 2.6) where the signals at $\delta_{\text{C}} = 204.6 \text{ ppm}$ and 207.4 ppm are observed as a doublet ($^2J_{\text{C–C}} = 15 \text{ Hz}$) and the four carbons of the cyclobutene are observed as doublet-of-doublet-of-doublets (ddd, Fig. S12 and 13, ESI[‡]). ATR-IR spectroscopic analysis of a mixture of **2** and **10** in Fomblin® gave the characteristic signal for **2** at $\nu(\text{CO}) = 1674 \text{ cm}^{-1}$ and three remaining strong stretches. The stretches at $\nu(\text{CO}) = 2027 \text{ cm}^{-1}$ and 1980 cm^{-1} are attributed to the metal



$\text{C}\equiv\text{O}$ groups and the signal at $\nu(\text{CO}) = 1732 \text{ cm}^{-1}$ to the ketone moiety of the cyclobutene. The cyclobutene $\text{C}=\text{O}$ stretch is again comparable to others in the literature.^{49,52,53}

While *in situ* IR and NMR data from phase 2 of the reaction show some evidence for iron–ketene intermediates, this is further supported by the isolation and structural characterisation of three complexes containing ketenyl moieties (Fig. 11). Complex **11** (Fig. 12) was isolated as extremely air sensitive orange crystals from a concentrated Et_2O solution of **4^{Mes}** in the absence of a CO atmosphere. Complex **11** features two ketenyl moieties bound to an iron centre, with one of the mesityl groups providing a stabilising η^3 interaction, which shows a comparable η^3 -distance to $[\text{Zn}(\mu\text{-Cl})(\text{C}_6\text{F}_5)(\eta^3\text{-C}_6\text{Me}_6)]_2$ reported by Bochmann *et al.*⁵⁴ The $\text{C}=\text{C}$ (C1–C2, 1.297(10) Å) and $\text{C}=\text{O}$ (C2–O1, 1.182(9) Å and C102–O8, 1.179(9) Å) bond distances in **11** are in accordance with the handful of iron–ketene complexes in the literature.^{14,55–72} So far, **11** has only been isolated once, and the crystallisation is challenging to replicate, so further characterisation of this species has not been possible.

Complexes **12** and **13** were isolated when performing the reaction in iso-hexane in the absence of an atmosphere of CO. **12** and **13** co-crystallised in the space group $P\bar{1}$ with one

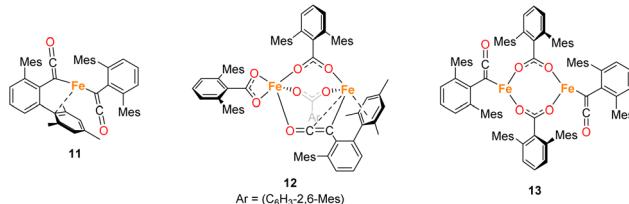


Fig. 11 Ketene containing structures **11**–**13** isolated during the course of these studies.

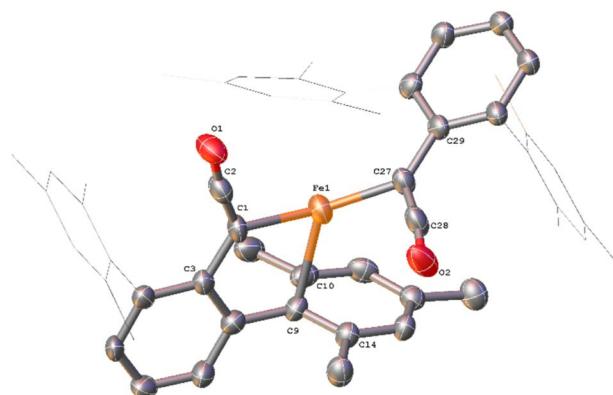


Fig. 12 View of the metal complex from the crystal structure of **11** with anisotropic displacement ellipsoids set at 50% probability. Mesityl groups shown as wireframe, apart from the one bonding to the Fe, and hydrogen atoms have been omitted for clarity. Selected bond distances (Å) and angles (°) for **11** shown: C1–C2 1.297(10), C2–O1 1.182(9), C27–C28 1.296(10), C28–O2 1.179(9), Fe1–C1 2.052(6), Fe1–C9 2.451(7), Fe1–C10 2.843(7), Fe1–C14 2.645(7), Fe1–C27 2.039(5), C1–C2–O1 171.6(7), C1–Fe1–C27 136.0(3), C27–C28–O2 172.2(6), Fe1–C1–C2 118.9(5), Fe1–C27–C29 127.9(5).

equivalent of **12** and half an equivalent of **13** per asymmetric unit (Fig. 13a). **12** is composed of one ketenyl and three carboxylate groups, with two of the carboxylates and the ketenyl moiety bridging two iron centres. The ketenyl binds to Fe2 through an η^2 -interaction with the C=C bond, and the remaining carboxylate group is bound in a terminal

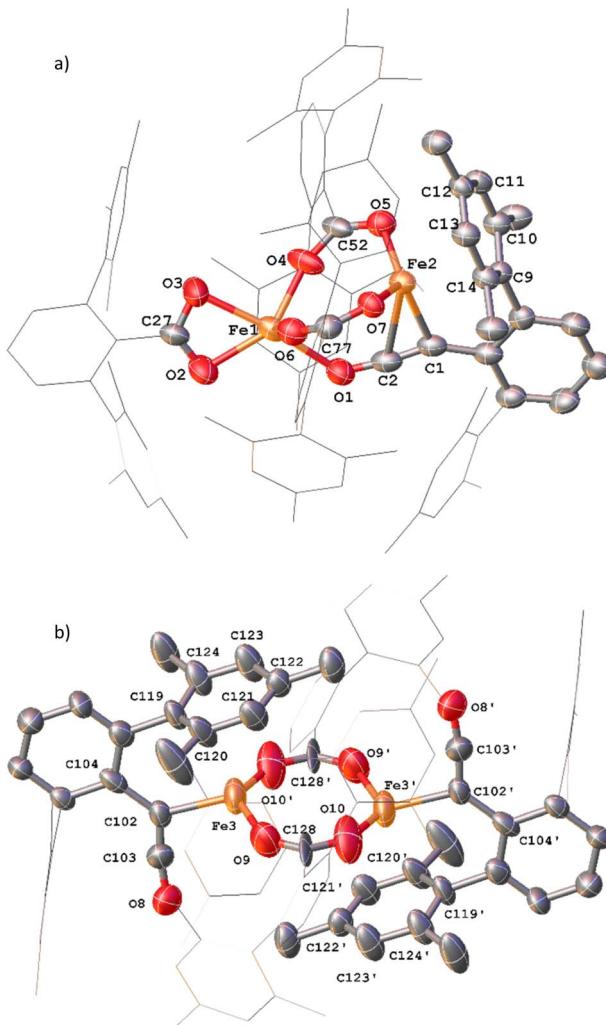


Fig. 13 (a) View of the metal complex from the crystal structure of **12** with anisotropic displacement ellipsoids set at 50% probability. 2,6-Mes₂C₆H₃ groups and non-bonding Mes groups shown as wireframe and hydrogen atoms have been omitted for clarity. Selected bond distances (Å) and angles (°) for **12** shown: C1–C2 1.254(9), C2–O1 1.233(8), Fe2–C1 2.133(6), Fe2–C2 2.475(7), Fe1–O1 2.169(5), Fe1–O2 2.104(5), Fe1–O3 2.224(5), Fe1–O4 2.016(4), Fe1–O6 2.024(5), Fe2–O5 1.987(4), Fe2–O7 1.984(4), C1–C2–O1 173.4(7), O2–C27–O3 118.1(6), O4–C52–O5 121.8(6), O6–C77–O7 124.9(6). (b) View of the metal complex from the crystal structure of **13** with anisotropic displacement ellipsoids set at 30% probability. Atoms marked with ' were generated using the following symmetry operator: 2 – X, –Y, 1 – Z. 2,6-Mes₂C₆H₃ groups and non-bonding Mes groups shown as wireframe and hydrogen atoms have been omitted for clarity. Selected bond distances (Å) and angles (°) for **13** shown: C102–C103 1.213(15), C103–O8 1.250(14), Fe3–C102 2.026(8), Fe3–O9 2.09(1), Fe3–O10' 2.018(11), C102–C103–O8 172.5(12), O9–C128–O10 115.6(11), Fe3–C102–C104 126.8(7).



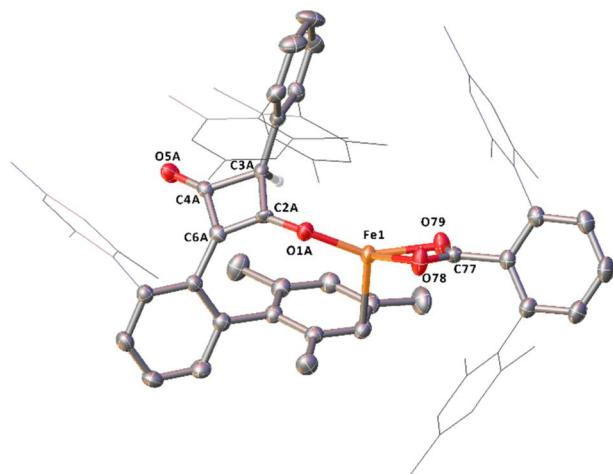


Fig. 14 Molecular structure of **14** with anisotropic displacement ellipsoids set at 50% probability, major component of disordered C_4 core shown. Mesityl groups shown as wireframe, except those interacting with the Fe and hydrogen atoms, except H_4 , have been omitted for clarity. Selected bond distances (\AA) and angles ($^\circ$) for **13** shown: Fe1–O1A 1.8702(9), Fe1–O78 2.1076(14), Fe1–O79 2.0442(10), C2A–O1A 1.306(17), C2A–C3A 1.560(17), C2A–C6A 1.362(18), C3A–C4A 1.576(13), C4A–C6A 1.456(17), C4A–O5A 1.20(2), O78–C77–O79 118.42(14).

coordination. The ketenyl $C=C$ ($C1-C2$, 1.254(9) \AA) and $C=O$ ($C2-O1$, 1.233(8) \AA) bond lengths are similar to those in **11**.

Complex **13** is situated across a special position in the asymmetric unit and contains two terminal ketenyl moieties and two bridging carboxylates (Fig. 13b). **13** shows comparable angles for the bridging carboxylates to **3** and **12**, but a slight elongation of one of the Fe–O bonds [Fe3–O10, 2.09(1) \AA]. For the ketene moiety in **13**, the $C=C$ bonds are shorter than the $C=O$, opposite to what is observed for **11** and **12**. The Fe–C bond for the ketenyl group [2.026(5) \AA] is also within error for a similar bond in **11**.

Complex **14**, which features a squaraine-like moiety bound to an Fe centre alongside a carboxylate ligand has also been identified by X-ray crystallography (Fig. 14). One of the mesityl groups of the squaraine-like moiety interacts in an η^3 fashion to the Fe with bond lengths comparable to **11**. The C_4 moiety in **14** shows significant asymmetry as demonstrated by the differing C–C and C–O distances and is disordered by inversion of the C_4 group. The C_4 cycle can be best described as a cyclobuteneone and is comparable to both Heimgartner's free cyclobuteneone and complex **10**.⁴⁹ The carboxylate moiety is comparable to that previously reported for the terminal carboxylate of **3**.

2.6 ^{13}C labelling experiments for phase 2

To further probe the mechanism, experiments with ^{13}CO were conducted. ^{13}CO labelled **4^{Mes}** (**4^{Mes-13C}**) was obtained from the reaction between **1^{Mes}** and ^{13}CO *via* the methodology described in Section 2.3. This was then reacted further under an atmosphere of either natural abundance CO (henceforth referred to as CO) or ^{13}CO . When reacting **4^{Mes-13C}** with ^{13}CO in $C_6\text{H}_6$ (see Section 4.2.10 of the ESI for details), to allow simultaneous IR

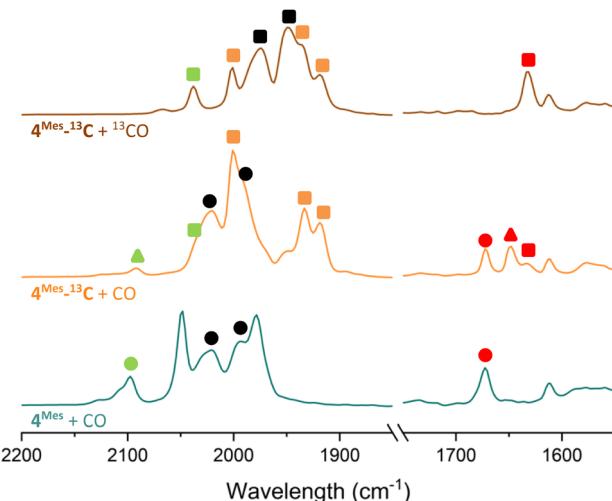


Fig. 15 Comparison of the IR spectra of reactions after 2 days of **4^{Mes}** and **4^{Mes-13C}** under an atmosphere of CO or ^{13}CO . All reactions performed in $C_6\text{H}_6$. Left, metal carbonyl and ketene region. Right, region for squaraine. Circles signify ^{12}C products, triangles signify mixed $^{12}\text{C}/^{13}\text{C}$ and square signifies ^{13}C products. Signals in green are for ketene (**8**). Signals in red are squaraine (**2^{Mes}**). Signals in black are $\text{Fe}(\text{CO})_5$.

and NMR spectroscopic analysis, the IR signals are red-shifted relative to the peaks for **4^{Mes}**. When **4^{Mes-13C}** was reacted under an atmosphere of CO, the IR spectrum recorded after 48 hours showed no evidence of CO exchange for **4^{Mes-13C}** as the signals for this species matched those observed for **4^{Mes-13C}** under an atmosphere of ^{13}CO (Fig. 15). However, the signals for **8** (from degradation processes) and $\text{Fe}(\text{CO})_5$ are blue-shifted relative to the spectra obtained using **4^{Mes-13C}** and ^{13}CO , suggesting de-enrichment. The signal for **8** is marginally red-shifted for **4^{Mes-13C}** + CO relative to **4^{Mes}** + CO ($\nu(\text{CO}) = 2091 \text{ cm}^{-1}$ vs. 2097 cm^{-1} , **8**. Cf. **8-13C** $\nu(\text{CO}) = 2037 \text{ cm}^{-1}$), indicative of an $\text{Ar-}^{13}\text{C} = ^{12}\text{C=O}$ ($\text{Ar} = 2,6\text{-MesC}_6\text{H}_3$) isotope pattern in **8** formed from **4^{Mes-13C}** + CO. Three IR signals are present for **2^{Mes}** when reacting **4^{Mes-13C}** under an atmosphere of natural abundance CO. This arises as the squaraine moiety has two C–O groups which can contain either ^{13}C or ^{12}C , and the three signals relate to squaraine C–O moieties containing $^{13}\text{C}/^{13}\text{C}$, $^{13}\text{C}/^{12}\text{C}$ or $^{12}\text{C}/^{12}\text{C}$. This suggests that uptake of CO is essential for the transformation of **4^{Mes}** into a ketenyl-containing species, as well as for the release of $\text{Fe}(\text{CO})_5$.

The $^{13}\text{C}\{^1\text{H}\}$ NMR spectra of reactions between **4^{Mes-13C}** and CO gives additional insight into the reaction pathway. Both compounds **8**, from decomposition, and **2^{Mes}** show uptake of carbon from the atmosphere, decreasing the degree of enrichment of the products (Fig. 16). For **8**, the signal at 193.1 ppm is observed solely as a doublet with an integral of 0.62 while the signal at 24.8 ppm splits into a doublet and a singlet with, a total integral of 1. Taking both the coupling pattern and integrations into account, this means that the C=O of the ketene ($\delta_{\text{C}} = 193.1 \text{ ppm}$) is partially incorporated from the atmosphere. Conversely, the C=C=O ($\delta_{\text{C}} = 24.8 \text{ ppm}$) is exclusively retained, presumably the carbene carbon in **4^{Mes-13C}**. This highlights the importance of an atmosphere of

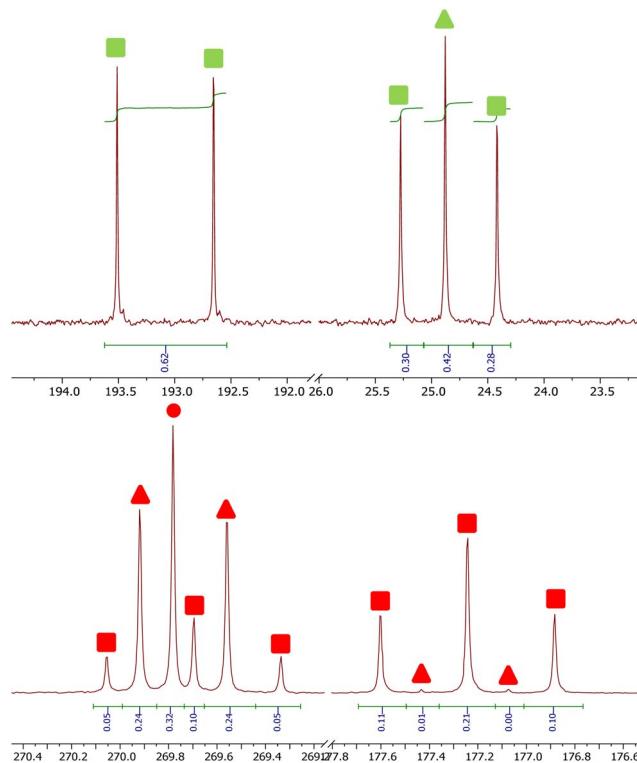


Fig. 16 ^{13}C containing products from the reaction between $4^{\text{Mes}}\text{-}^{13}\text{C}$ and CO. Top, signals for **8**. Bottom, signals for 2^{Mes} . Circles signify ^{12}C products, triangles signify mixed $^{12}\text{C}/^{13}\text{C}$ and square signifies ^{13}C products. Markers in green are for ketene (**8**) and markers in red are squaraine (2^{Mes}).

CO in the formation of the ketenyl complex. Furthermore, performing the same analysis for 2^{Mes} indicates that there is a 56% depletion for the C=O ($\delta_{\text{C}} = 177.2$ ppm, integral = 0.44) relative to the $(2,6\text{-Mes}_2\text{C}_6\text{H}_3)\text{-C}$ ($\delta_{\text{C}} = 269.8$ ppm, integral = 1). This is confirmed through analysis of the coupling patterns where the $(2,6\text{-Mes}_2\text{C}_6\text{H}_3)\text{-C}$ shows a ratio of 20 (t) : 48 (d) : 32 (s), consistent with 44% of the neighbouring carbons being ^{13}C enriched. For the C=O of 2^{Mes} a ratio of 98 (t) : 2 (d), consistent with 98% carbons at the neighbouring positions being enriched, which is expected as the ^{13}CO used is 99% enriched. This again shows that an atmosphere of CO is essential for the reaction to progress cleanly. It also suggests that there may be more than one step where the CO is incorporated due to the difference between the enrichment of **8** and 2^{Mes} .

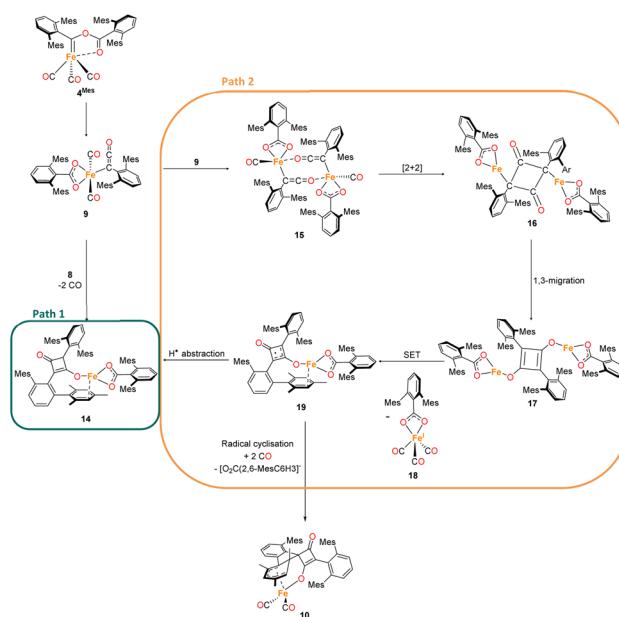
The EPR studies (Section 2.4) indicate that radicals are present in the reactions. While one of the signals observed correlates to $[2]^{+}$, the signal at $g_{\text{iso}} = 2.038$ has still not been assigned. When enriching the sample with ^{13}CO , simulations of the experimental spectrum as a $^{13}\text{C}_4$ -core using two pairs of ^{13}C couplings, as is the case for $[2]^{+}$, were unconvincing. A better simulation was obtained by considering lower symmetry in a $^{13}\text{C}_4$ -core with one larger coupling on a single ^{13}C environment and a smaller coupling across three ^{13}C atoms (Fig. S68, 69 and Table S3, ESI†). This splitting is consistent with that expected for a radical centered on a species similar to a squaraine where the C_4 core bears both a C=O and C–O group with

delocalisation of the radical over three carbon atoms. For reactions between CO and $4^{\text{Mes}}\text{-}^{13}\text{C}$, the EPR signal obtained for $[2]^{+}$ shows depletion of the ^{13}C for the C=O, and the signal was modelled with a 43% ^{13}C enrichment. This is consistent with that observed in the $^{13}\text{C}\{^1\text{H}\}$ NMR spectra. The EPR studies show that the environment in which the reaction proceeds is highly reducing, with $[2]^{+}$ forming in the absence of an external reductant. Thus, single electron reactions are highly plausible.

2.7 Proposed mechanism for phase 2

The mechanism for phase 2 of the reaction (4^{Mes} to 2^{Mes} , 3^{Mes} , $\text{Fe}(\text{CO})_5$) is more tentative than phase 1, but we propose a plausible route based on the data available. Some key observations are that ketene **8** and an iron-ketene complex (**9**) are observed spectroscopically (*in situ* IR and NMR) and that three iron-ketene containing byproducts (**11–13**) were isolated from the reaction. Thus, an iron-ketene complex is likely a key intermediate. We must also account for the formation of cyclobutanones **10** and **14** (Fig. 10 and 14), which are structurally related to squaraine 2^{Mes} . We therefore propose that 4^{Mes} initially rearranges to an iron ketene/carboxylate **9** (Scheme 4). While **9** has not been definitively characterised, its formation is consistent with the observed data, and we tentatively assign the iron-ketene signals (Fig. 7 and 8) to **9**.

From **9**, two possible paths to the isolated cyclobutanone complex **14** can be envisaged (Scheme 4). Path 1 involves a $[2 + 2]$ cycloaddition between **9** and **8**, followed by a 1,3-migration of the Fe centre, leading directly to **14**. However, this route would not explain the presence of radical species detected by EPR



Scheme 4 Two potential routes to **14** via proposed complex **9**. Path 1 forms **14** via a $[2 + 2]$ cycloaddition of **8** and **9**. Path 2 forms **14** through dimerization of **9** then a $[2 + 2]$ cyclisation, followed by 1,3 migration and a single electron transfer, giving **18** and **19**. Path 2 also has the potential to form decomposition complex **10** via radical cyclisation of **19**.



spectroscopy. We therefore suggest that Path 2, which involves Single Electron Transfer (SET) processes, is more plausible (Scheme 4). In path 2, two molecules of **9** dimerise to form **15**. **15** then undergoes an intramolecular [2 + 2] cycloaddition to form **16** followed by two 1,3-migrations to form **17**, reducing steric crowding. From **17**, an Fe(i) complex (**18**) and an Fe complex with a ligand centred radical (**19**) are formed by SET. Cyclobutene complexes **10** and **14** can both be formed from the proposed **19**, either by radical cyclisation (**10**) or H-atom abstraction (**14**).

From this, we propose the overall mechanism shown in Scheme 5. Here, the reaction proceeds as in Scheme 4 up to the formation of the ligand-centred radical **19**. This undergoes an additional SET to form the squaraine **2^{Mes}**, along with another equivalent of Fe(i) species **18**. We then propose that **18** undergoes disproportionation to form the other major products, iron carboxylate **3^{Mes}** and $\text{Fe}(\text{CO})_5$. Since it was not possible to observe Fe(i) signals by EPR or MB spectroscopy, we suggest the disproportionation of **18** is rapid. This is consistent with the formation of $\text{Fe}(\text{CO})_5$ in solution, and precipitation of **3^{Mes}** within 24 hours. Excess CO is required for clean formation of $\text{Fe}(\text{CO})_5$, otherwise the reactive intermediates undergo alternative reactions, resulting in cyclobutene **10** and the unknown Fe complexes detected by MB spectroscopy (**A** and **B**, Fig. 9). It is worth noting that, **A** and **B** are always observed by MB spectroscopy due to the high concentrations required for this technique, meaning that it is not possible for excess CO in solution to react for **4^{Mes}** to react cleanly.

Further DFT studies for phase 2 of the reaction proved difficult due to the size of the dimer like molecules. After the

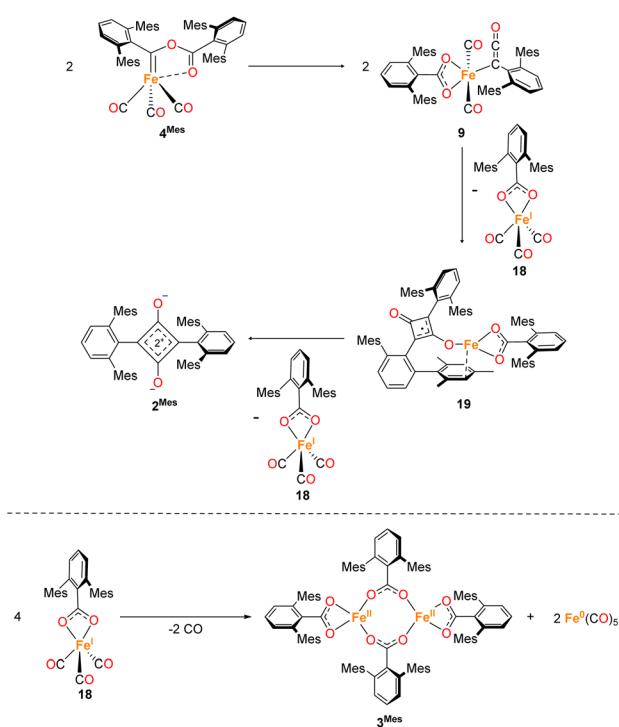
formation of **4^{Mes}**, high-spin Fe is strongly preferred according to the calculations. This is consistent with the observed formation of paramagnetic species in the ^1H NMR spectra (see Fig. S61†).

3. Conclusions

Through combined spectroscopic and structural studies we have developed a plausible mechanism for the formation of a squaraine (**2^{Mes}**) from two-coordinate **1^{Mes}**, *via* reaction with CO. The first part of reaction clearly shows the formation of diamagnetic complexes by ^1H NMR spectroscopy and a number of carbonyl containing complexes with complexes **5–7** isolated en route to the stable complex **4^{Mes}**, which are envisaged to form though sequential migratory insertion reactions. The reactivity post formation of **4^{Mes}** is significantly more difficult to follow due to the high sensitivity and reactivity of the intermediate species. Our initial studies proposed that Fe–ketenyl complexes were responsible for the formation of **2** due to a characteristic signal at 2097 cm^{-1} in the IR spectrum, however, this signal was found to relate to a protonated ketene (**8**) which forms in the presence of silicon grease and/or moisture. Other signals which are proposed to relate to Fe–ketenyl complexes are fleetingly observed by IR spectroscopy, but support them being key intermediates to the formation of **2**. Pleasingly, other ketene complexes were isolated over the course of these studies (**11–13**) which gives strong evidence that ketenyl complexes are present en route to **2**. However, these are proposed to be decomposition products and not active in the formation of **2**. Complexes **10** and **14** both contain ligands bearing cyclobuteneone moieties, providing further clues on potential intermediates in this reaction, all of which are derived from ketenes/ketenyl groups. Further spectroscopic studies using labelled materials and performing the reaction in the absence of an atmosphere of CO highlight the importance of excess CO in the reaction. Reactions of labelled **4^{Mes}-¹³C** under natural abundance showed selective loss of ^{13}CO from the C–O of the squaraine product indicating that key steps post **4^{Mes}** forming ketenes requires uptake of CO from solution. The enriched carbene C in **4^{Mes}-¹³C** is retained adjacent to the terphenyl moiety in **2^{Mes}**. EPR spectra at room temperature showed evidence for the formation of radicals which are consumed over the course of the reaction, however, they could not be unequivocally identified. Mössbauer spectroscopy showed that the iron intermediates post **4^{Mes}** react rapidly, forming $\text{Fe}(\text{CO})_5$ within 24 hours. While there is no direct evidence for an Fe(i) complex by EPR and MB spectroscopy, it is possible to propose a series of SET reactions which form a fleeting Fe(i) complex (**18**) that undergoes rapid disproportionation to $\text{Fe}(\text{CO})_5$ and **3^{Mes}**. This work demonstrates how a multi-spectroscopic and structural approach is needed to truly understand highly complex reactions, showcasing the benefits of both *in situ* measurements and reaction sampling methodology.

Data availability

A data repository containing spectroscopic data in their raw (IR and NMR) and processed forms (mass spectrometry, CHN, EPR



Scheme 5 Proposed formation of **2^{Mes}** and **13** via a ketenyl complex (**9**) by sequential single electron transfers, forming Fe(i) complex **18**. **18** then undergoes disproportionation resulting in **3^{Mes}** and $\text{Fe}(\text{CO})_5$.



and Mössbauer spectroscopy) can be found *via* the following <https://doi.org/10.17639/nott.7407>.

Author contributions

NTC: drafted the manuscript. DLK, JM, NTC and LJT: conceptualization and supervision of the project. NTC, LJT, NC, YL, DR, MLN and DLK: writing, reviewing and editing of the manuscript. NTC, DR, NC: compiled the ESI. DR, AB, LJT: performed DFT analysis. YL, NTC, SPA: collected XRD data in house, solved and refined the crystal structures. NTC and SPA: performed XRD data collections at Diamond Light Source and validated the refined XRD data. ESD: collection and processing of EPR data. PJM: assistance with ReactIR methodology and setup. KB: implemented bespoke and high resolution NMR data collection. BP-G: performed MALDI mass spectrometry. MLN, NC: performed Mössbauer measurements and analysis. NTC, YL, NC: contributed to the experimental work, formal analysis and data curation.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Dr Emma Richards, Cardiff University, is thanked for valuable discussions on EPR experiments. DLK, JM, DR and NTC thank the Leverhulme Trust (RPG-2021-183) for funding. YL and DLK thank the University of Nottingham for a Vice-Chancellor's Scholarship for Research Excellence. DR thanks Nottingham Trent University for provision of the Hamilton and Avicenna high-performance computing clusters, on which the calculations were performed. The authors would like to thank Diamond Light Source for beamtime (proposal CY28766), and the staff of beamline I19 for assistance with crystal testing and data collection.

References

- 1 M. E. Dry, *Appl. Catal., A*, 1996, **138**, 319–344.
- 2 A. Y. Khodakov, W. Chu and P. Fongarland, *Chem. Rev.*, 2007, **107**, 1692–1744.
- 3 G. V. Schulz, *Z. Phys. Chem.*, 1935, **30B**, 379–398.
- 4 P. J. Flory, *J. Am. Chem. Soc.*, 1936, **58**, 1877–1885.
- 5 R. A. Friedel and R. B. Anderson, *J. Am. Chem. Soc.*, 1950, **72**, 2307.
- 6 L. Gmelin, *Ann. Phys.*, 1825, **80**, 31–62.
- 7 M. D. Anker, M. S. Hill, J. P. Lowe and M. F. Mahon, *Angew. Chem., Int. Ed.*, 2015, **54**, 10009–10011.
- 8 R. Lalrempuia, C. E. Kefalidis, S. J. Bonyhady, B. Schwarze, L. Maron, A. Stasch and C. Jones, *J. Am. Chem. Soc.*, 2015, **137**, 8944–8947.
- 9 K. Yuvaraj, I. Douair, A. Paparo, L. Maron and C. Jones, *J. Am. Chem. Soc.*, 2019, **141**, 8764–8768.
- 10 A. Paparo, K. Yuvaraj, A. J. R. Matthews, I. Douair, L. Maron and C. Jones, *Angew. Chem., Int. Ed.*, 2021, **60**, 630–634.
- 11 W. Yang, A. White and M. R. Crimmin, *Angew. Chem., Int. Ed.*, 2024, **63**, e202319626.
- 12 H. Braunschweig, T. Dellermann, R. D. Dewhurst, W. C. Ewing, K. Hammond, J. O. C. Jimenez-Halla, T. Kramer, I. Krummenacher, J. Mies, A. K. Phukan and A. Vargas, *Nat. Chem.*, 2013, **5**, 1025–1028.
- 13 A. Heilmann, J. Hicks, P. Vasko, J. M. Goicoechea and S. Aldridge, *Angew. Chem., Int. Ed.*, 2020, **59**, 4897–4901.
- 14 M. Majumdar, I. Omlor, C. B. Yildiz, A. Azizoglu, V. Huch and D. Scheschke, *Angew. Chem., Int. Ed.*, 2015, **54**, 8746–8750.
- 15 X. Wang, Z. Zhu, Y. Peng, H. Lei, J. C. Fettinger and P. P. Power, *J. Am. Chem. Soc.*, 2009, **131**, 6912–6913.
- 16 M. Xu, A. R. Jupp and D. W. Stephan, *Angew. Chem., Int. Ed.*, 2019, **58**, 3548–3552.
- 17 T. Matsuo and H. Kawaguchi, *J. Am. Chem. Soc.*, 2005, **127**, 17198–17199.
- 18 S. Hu, T. Shima and Z. Hou, *J. Am. Chem. Soc.*, 2020, **142**, 19889–19894.
- 19 T. Watanabe, Y. Ishida, T. Matsuo and H. Kawaguchi, *J. Am. Chem. Soc.*, 2009, **131**, 3474–3475.
- 20 T. Kurogi, Y. Ishida, T. Hatanaka and H. Kawaguchi, *Dalton Trans.*, 2013, **42**, 7510–7513.
- 21 J. A. Buss and T. Agapie, *Nature*, 2016, **529**, 72–75.
- 22 J. A. Buss and T. Agapie, *J. Am. Chem. Soc.*, 2016, **138**, 16466–16477.
- 23 R. Yadav, T. Simler, M. T. Gamer, R. Köppe and P. W. Roesky, *Chem. Commun.*, 2019, **55**, 5765–5768.
- 24 D. H. Berry, J. E. Bercaw, A. J. Jircitano and K. B. Mertes, *J. Am. Chem. Soc.*, 1982, **104**, 4712–4715.
- 25 A. S. Batsanov, J. A. Cabeza, M. G. Crestani, M. R. Fructos, P. García-Álvarez, M. Gille, Z. Lin and T. B. Marder, *Angew. Chem., Int. Ed.*, 2016, **55**, 4707–4710.
- 26 T. Shima and Z. Hou, *J. Am. Chem. Soc.*, 2006, **128**, 8124–8125.
- 27 W. J. Evans, A. L. Wayda, W. E. Hunter and J. L. Atwood, *J. Chem. Soc., Chem. Commun.*, 1981, 706–708.
- 28 W. J. Evans, J. W. Grate, L. A. Hughes, H. Zhang and J. L. Atwood, *J. Am. Chem. Soc.*, 1985, **107**, 3728–3730.
- 29 W. J. Evans, D. S. Lee, J. W. Ziller and N. Kaltsoyannis, *J. Am. Chem. Soc.*, 2006, **128**, 14176–14184.
- 30 P. L. Arnold, Z. R. Turner, R. M. Bellabarba and R. P. Tooze, *Chem. Sci.*, 2010, **2**, 77–79.
- 31 B. M. Gardner, J. C. Stewart, A. L. Davis, J. McMaster, W. Lewis, A. J. Blake and S. T. Liddle, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 9265–9270.
- 32 A. M. Geer, D. L. Kays and L. J. Taylor, in *Homologation Reactions*, John Wiley & Sons, Ltd, 2023, pp. 813–846.
- 33 R. Y. Kong and M. R. Crimmin, *Dalton Trans.*, 2020, **49**, 16587–16597.
- 34 C. Ni and P. P. Power, *Chem. Commun.*, 2009, 5543–5545.
- 35 H. R. Sharpe, A. M. Geer, L. J. Taylor, B. M. Gridley, T. J. Blundell, A. J. Blake, E. S. Davies, W. Lewis, J. McMaster, D. Robinson and D. L. Kays, *Nat. Commun.*, 2018, **9**, 3757.
- 36 G. Doyle, K. A. Eriksen and D. Van Engen, *J. Am. Chem. Soc.*, 1985, **107**, 7914–7920.



37 L. S. Crocker, B. M. Mattson, D. M. Heinekey and G. K. Schulte, *Inorg. Chem.*, 1988, **27**, 3722–3725.

38 P. A. Agron, R. D. Ellison and H. A. Levy, *Acta Crystallogr., Sect. C*, 1991, **47**, 913–916.

39 X. Tao, M. Feng, Y.-Y. Zhang, Y.-Q. Li, N. Wang and Y.-Z. Shen, *Chin. J. Struct. Chem.*, 2009, **28**, 287–290.

40 H. Braunschweig, R. D. Dewhurst and C. Schneider, *Organometallics*, 2016, **35**, 1002–1007.

41 D. L. Kays (née Coombs) and A. R. Cowley, *Chem. Commun.*, 2007, 1053–1055.

42 L. M. Epstein, *J. Chem. Phys.*, 1962, **36**, 2731–2737.

43 D. J. Evans, M. Jimenez-Tenorio and G. J. Leigh, *J. Chem. Soc., Dalton Trans.*, 1991, 1785–1787.

44 P. Gülich, E. Bill and A. X. Trautwein, *Mössbauer Spectroscopy and Transition Metal Chemistry*, Springer, Berlin, Heidelberg, 2011.

45 T. S. Cameron, M. D. Clerk, A. Linden, K. C. Sturge and M. J. Zaworotko, *Organometallics*, 1988, **7**, 2571–2573.

46 R. S. Hay-Motherwell, B. Hussain-Bates, M. B. Hursthouse, B. E. Mann and G. Wilkinson, *J. Chem. Soc., Dalton Trans.*, 1993, 3219–3228.

47 W. B. Wulff, A. L. Rheingold, and M. B. Allen, *CSD Communication, Structure code: HIGREU*.

48 E. A. Trifonova, D. S. Perekalin, N. L. Loskutova, Y. V. Nelyubina and A. R. Kudinov, *J. Organomet. Chem.*, 2014, **770**, 1–5.

49 L. S. Trifonov, A. S. Orahovats, R. Prewo, J. H. Bieri and H. Heimgartner, *J. Chem. Soc., Chem. Commun.*, 1986, 708–709.

50 P. T. DiMauro and P. T. Wolczanski, *Organometallics*, 1987, **6**, 1947–1954.

51 R. U. Kirss, A. Henriksen, D. A. Forsyth and W. Feighery, *Inorg. Chim. Acta*, 2006, **359**, 4393–4397.

52 L. D. Foland, J. O. Karlsson, S. T. Perri, R. Schwabe, S. L. Xu, S. Patil and H. W. Moore, *J. Am. Chem. Soc.*, 1989, **111**, 975–989.

53 G. Verniest and N. D. Kimpe, *Synlett*, 2005, **2005**, 0947–0950.

54 Y. Sarazin, J. A. Wright, D. A. J. Harding, E. Martin, T. J. Woodman, D. L. Hughes and M. Bochmann, *J. Organomet. Chem.*, 2008, **693**, 1494–1501.

55 J. Martin-Gil, J. A. K. Howard, R. Navarro and F. G. A. Stone, *J. Chem. Soc., Chem. Commun.*, 1979, 1168–1169.

56 J. W. Kolis, E. M. Holt and D. F. Shriver, *J. Am. Chem. Soc.*, 1983, **105**, 7307–7313.

57 S. Ching, E. M. Holt, J. W. Kolis and D. F. Shriver, *Organometallics*, 1988, **7**, 892–898.

58 P. L. Bogdan, M. Sabat, S. A. Sunshine, C. Woodcock and D. F. Shriver, *Inorg. Chem.*, 1988, **27**, 1904–1910.

59 Y. Tang, J. Sun and J. Chen, *J. Chem. Soc., Dalton Trans.*, 1998, 4003–4008.

60 H. Wadeohl, U. Arnold, H. Pritzkow, M. J. Calhorda and L. F. Veiros, *J. Organomet. Chem.*, 1999, **587**, 233–243.

61 R. Bertani, M. Casarin, P. Ganis, C. Maccato, L. Pandolfo, A. Venzo, A. Vittadini and L. Zanotto, *Organometallics*, 2000, **19**, 1373–1383.

62 V. V. Burlakov, P. Arndt, W. Baumann, A. Spannenberg and U. Rosenthal, *Organometallics*, 2006, **25**, 1317–1320.

63 M. A. Alvarez, M. E. García, M. E. Martínez, S. Menéndez and M. A. Ruiz, *Organometallics*, 2010, **29**, 710–713.

64 M. Landman, B. van der Westhuizen, D. I. Bezuidenhout and D. C. Liles, *Acta Crystallogr., Sect. E*, 2012, **68**, m308–m309.

65 H.-J. Liu, M. S. Ziegler and T. D. Tilley, *Angew. Chem., Int. Ed.*, 2015, **54**, 6622–6626.

66 S. A. Gonsales, I. Ghiviriga, K. A. Abboud and A. S. Veige, *Dalton Trans.*, 2016, **45**, 15783–15785.

67 V. Jakhar, D. Pal, I. Ghiviriga, K. A. Abboud, D. W. Lester, B. S. Sumerlin and A. S. Veige, *J. Am. Chem. Soc.*, 2021, **143**, 1235–1246.

68 W. Feuerstein, P. Varava, F. Fadaei-Tirani, R. Scopelliti and K. Severin, *Chem. Commun.*, 2021, **57**, 11509–11512.

69 T. Simler, K. N. McCabe, L. Maron and G. Nocton, *Chem. Sci.*, 2022, **13**, 7449–7461.

70 R. J. Ward, I. del Rosal, S. P. Kelley, L. Maron and J. R. Walensky, *Chem. Sci.*, 2023, **14**, 2024–2032.

71 R. Wei, X.-F. Wang, D. A. Ruiz and L. L. Liu, *Angew. Chem., Int. Ed.*, 2023, **62**, e202219211.

72 K. Yuvaraj, J. C. Mullins, T. Rajeshkumar, I. Douair, L. Maron and C. Jones, *Chem. Sci.*, 2023, **14**, 5188–5195.

