

Cutting the cost of carbon capture: a case for carbon capture and utilization†

Lennart Joos,^a Johanna M. Huck,^b Veronique Van Speybroeck^a and Berend Smit^{*bcd}

Received 29th February 2016, Accepted 11th March 2016

DOI: 10.1039/c6fd00031b

A significant part of the cost for carbon capture and storage (CCS) is related to the compression of captured CO₂ to its supercritical state, at 150 bar and typically 99% purity. These stringent conditions may however not always be necessary for specific cases of carbon capture and utilization (CCU). In this manuscript, we investigate how much the parasitic energy of an adsorbent-based carbon capture process may be lowered by utilizing CO₂ at 1 bar and adapting the final purity requirement for CO₂ from 99% to 70% or 50%. We compare different CO₂ sources: the flue gases of coal-fired or natural gas-fired power plants and ambient air. We evaluate the carbon capture performance of over 60 nanoporous materials and determine the influence of the initial and final CO₂ purity on the parasitic energy of the carbon capture process. Moreover, we demonstrate the underlying principles of the parasitic energy minimization in more detail using the commercially available NaX zeolite. Finally, the calculated utilization cost of CO₂ is compared with the reported prices for CO₂ and published costs for CCS.

1 Introduction

Nowadays, there are no technological hurdles to the industrial deployment of post-combustion carbon capture. However, the high energy penalty of the process and the associated financial costs hamper the actual use of the technology. Drastically lowering the price of this process is therefore a requisite for the competitiveness of carbon capture as a viable technology to reduce CO₂ emissions.^{1–5}

^aCenter for Molecular Modeling, Ghent University, B-9052 Zwijnaarde, Belgium. E-mail: veronique.vanspeybroeck@ugent.be

^bDepartment of Chemical and Biomolecular Engineering, University of California, Berkeley, CA 94720-1462, USA

^cDepartment of Chemistry, University of California, Berkeley, CA 94720-1462, USA

^dLaboratory of Molecular Simulation, Institut des Sciences et Ingénierie Chimiques, Valais, Rue d'Industrie 17, Ecole Polytechnique Fédérale de Lausanne (EPFL), CH-1950 Sion, Switzerland. E-mail: berend.smit@epfl.ch

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c6fd00031b



The most mature technology for carbon capture and storage (CCS) to separate CO₂ from the post-combustion flue gases is amine scrubbing, a two-step process in which CO₂ is first chemically bound to the amines and later recovered in a pure form by 'stripping' the CO₂ from the amines at high temperatures.^{6,7} A commonly used amine in this process is monoethanolamine (MEA). Heating the dilute aqueous amine mixture imposes a severe energy penalty on the process. Alternative technologies are therefore proposed, such as adsorption in nanoporous materials. A variety of classes of nanoporous materials exists with high CO₂/N₂ selectivity, good CO₂ uptake, as well as a less energy demanding regeneration than the amine solutions.^{8–10} Among them are metal–organic frameworks (MOFs),^{11,12} zeolitic imidazolate frameworks (ZIFs),^{13,14} porous polymer networks (PPNs),^{15–17} and zeolites.^{18–21}

There are several possibilities to lower the cost of post-combustion carbon capture with nanoporous materials, which will be outlined in the following subsections. In the first place, ongoing advances in material research and computational screening methods can help to design a new material with potentially better CCS properties. Secondly, novel ways to operate the carbon capture process can also reduce the energy consumption. Thirdly, changing the operating conditions of the carbon capture process also provides opportunities to save energy. The latter will be the focus of this manuscript.

1.1 Alternative materials

Due to rational material synthesis approaches there has been an explosion of new materials in recent years. Since their discovery in the 1990's, some 10 000 Metal–Organic Frameworks (MOFs) have been synthesized, and countless new ones as well as specific modifications are designed every year.²³ Moreover, the importance of zeolites in the petrochemical industry has spurred new developments in the field of zeolite synthesis as well.^{24,25}

To compare nanoporous materials among each other, Bae and Snurr presented five criteria to evaluate CCS performance: CO₂ uptake, CO₂/N₂ selectivity, CO₂ working capacity, regenerability of the material and a sorbent selection parameter.²⁶ Some of the authors of this manuscript proposed to use one metric instead, the "parasitic energy", *i.e.* the energy output of the power plant immediately consumed by the CCS process.^{27,28} This metric can easily be related to the industrial application of CCS, as it straightforwardly identifies the material with the lowest energy requirement for CCS.

However, experimentally determining any CCS metric for thousands of possible candidates would be nearly impossible. Therefore, high-throughput computational screenings of databases with existing as well as hypothetical materials have become very attractive tools. They allow us to evaluate millions of materials for their CCS performance and guide experimental efforts in the direction of the most promising materials. With these computer screenings, it is not only possible to assess the boundaries of the material space for their carbon capture performance, but also get molecular insights on the best performing materials. A screening of nanoporous materials from different material families identified that adsorbent-based carbon capture can yield a parasitic energy 30% lower than that of the amine scrubbing process, but that further improvements are not to be expected.^{27,28} Finally, going beyond brute force computational



screenings requires machine learning techniques, which have already been applied for gas adsorption in nanoporous materials.^{12,29,30}

1.2 Alternative technologies

Although the discovery of new materials may gradually improve carbon capture performance, previously mentioned screening studies showed that this approach has its limits. Therefore, it is worth taking a second look at the carbon capture process, shown in Fig. 1. In conventional nanoporous materials-based carbon capture process, the exhaust gases of the power plant are first cooled and CO₂ is adsorbed at low temperatures to maximize CO₂ uptake. Once the adsorption bed is saturated with CO₂, the nanoporous material is regenerated either by heating the material or by applying a vacuum. Either way, a parasitic energy penalty is imposed on the process. Moreover, H₂O in the exhaust gases will compete for the same adsorption sites as CO₂, thereby lowering the CO₂ uptake and increasing the parasitic energy.³¹

In the newly proposed “High-temperature Adsorption and Low-temperature Desorption” (HALD) set-up, the temperature-dependent competitive adsorption of CO₂ and H₂O is exploited to overcome the high energy requirement.^{22,32} At high temperatures, the competitive adsorption of CO₂ and H₂O is in favor of CO₂, so adsorption at higher temperatures (without first cooling the exhaust gases) improves the selectivity towards CO₂. As the competition is in favor of H₂O at low temperatures, CO₂ can be desorbed at low temperatures, by cooling and saturating the adsorbent with H₂O. When the regenerated zeolite – void of CO₂ but full with H₂O – is brought into contact again with the hot exhaust gases, the waste energy of the flue gas desorbs the H₂O from the material and restores the

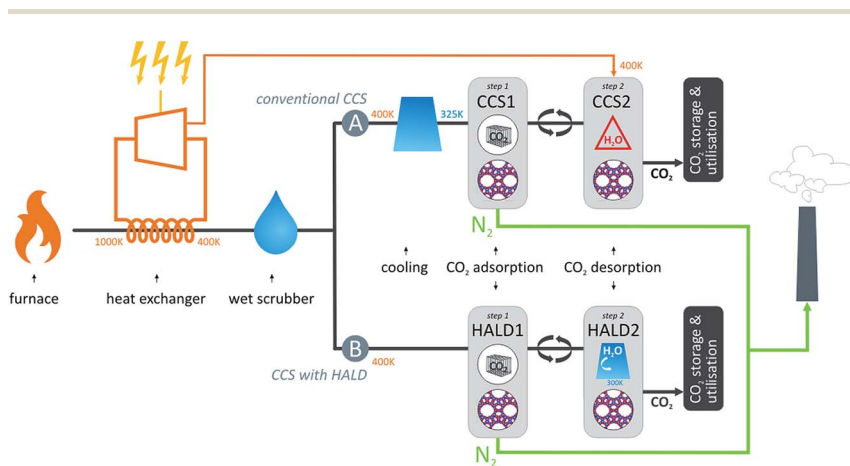


Fig. 1 Layout of a coal-fired power plant. (A) represents a power plant retrofitted with conventional carbon capture technology based on CO₂ adsorption on a nanoporous material. The exhaust gases are first cooled and CO₂ is subsequently adsorbed. Regeneration of the bed requires a parasitic energy from the power plant, which is even higher when H₂O is present. (B) is a power plant retrofitted with the proposed High-temperature Adsorption and Low-temperature Desorption (HALD) technology, whereby adsorption takes place at high temperatures and the material is regenerated by cooling the adsorbent while saturating it with H₂O.²²



competitive advantage of CO₂. This temperature swing process utilizes the residual heat of the exhaust gases to overcome the competition with H₂O and capture the CO₂ from this CO₂/N₂/H₂O mixture and hence does not require the input of energy for regeneration. The upside-down HALD alternative to traditional CCS methods thereby opens new perspectives to reduce the energy penalty of CCS.

1.3 Alternative operating conditions

For the application of carbon capture and storage (CCS), the captured CO₂ has to be of high purity (99%) before it can be compressed to a supercritical state at around 150 bar.³³ Only in these conditions can CO₂ efficiently be stored in geological formations. The high purity requirement and the compression to high pressure constitute a large part of the energy penalty for CCS. However, the stringent conditions may not always be necessary for specific cases of carbon capture and utilization (CCU). As a result, the price per tonne CO₂ may actually be significantly lower, shifting the economics of CCU into a more favorable direction.

In Table 1, a few examples of CCU are given. From left to right, the examples are organized with decreasing CO₂ purity for the application, 99%, 70% or 50%. From top to bottom, the CO₂ source is indicated: coal-fired power plants, with 14% CO₂ in their flue gases, gas turbine exhaust gases, with about 4% CO₂ flue gases, and air, with CO₂ levels of about 400 ppm.

An application in which 50% CO₂ could suffice, but for which large amounts of CO₂ are needed, is the cultivation of algae for the production of biofuels.³⁴ Increased CO₂ levels are used to heighten the crop production in greenhouses, although for this application smaller quantities suffice, provided by a gas turbine or a biomass plant for instance.^{35,36} Finally, for the carbonation of drinks, local CO₂ production directly from the air could be an option. An interesting example for 70% CO₂ is the production of "Carbstone".³⁷ In a chemical process, finely ground CaO and MgO, waste products from the metallurgical industry, react with CO₂ to produce CaCO₃/MgCO₃ bricks and tiles. This is an example where CO₂ is both permanently sequestered in a chemically stable solid product, as well as utilized for a useful application (CCUS). Smaller amounts of CO₂ could be used as an inert gas in welding,³⁸ or for specific molding techniques, in which CO₂ was found to increase the hardness.³⁹ Finally, locally producing small amounts of CO₂ could also be practical for fumigation, for instance of bed bugs.⁴⁰ Some cases however do require the use of very pure, 99% CO₂. Enhanced Oil Recovery (EOR) for instance, injects CO₂ in near-depleted oil wells under supercritical conditions.⁴¹ The market for this application is in the order of millions of tonnes of CO₂ per year, and therefore large coal-fired power plants are needed to fill that

Table 1 Examples of carbon capture and utilization, organized according to the CO₂ content of the CO₂ source (in) and the desired CO₂ content for the application (out)

In/out	99%	70%	50%
COAL	EOR	Carbstone	Algae
NG	Fire extinguisher	Welding/molding	Greenhouses
AIR	CCS	Fumigation	Bubbles in drinks



demand. For fire extinguishers, pure CO₂ is also necessary, but much smaller CO₂ sources can be utilized.⁴² Finally, direct air capture could serve as a solution of last resort, when the climate turns really bad and the CO₂ level in the atmosphere has to be suppressed on a very short term, direct air capture devices can be used for carbon capture and storage, directly from the atmosphere.^{43,44}

2 Goal

The central question in this manuscript is how much the parasitic energy of carbon capture can be lowered by utilizing CO₂ at 1 bar and adapting the final purity requirements for CO₂ from different CO₂ sources. First, we investigate the influence of the final purity requirement for the case of coal-fired power plant flue gases. Then, we extend this analysis to other sources of CO₂, the flue gases of natural gas-fired power plants and ambient air. Afterwards, we go into more detail on the purity of the desorbed mixture. Then, the properties of NaX, a commercially available material, and the minimization of the parasitic energy are studied in more detail. Finally, we will compare the calculated CCU cost with reported costs for CCS and available prices for CO₂.

3 Methodology

The parasitic energy of the carbon capture process stems from the regeneration of the nanoporous material, either by heating the material, by applying a vacuum or a combination of the two. In the first place, the material can be heated up in order to trigger CO₂ desorption. This thermal heating energy, Q_{thermal} , not only incorporates the sensible heat requirement to heat the material but also the required desorption heat of CO₂. The latter heat term is required to undo the binding of the CO₂ to the nanoporous material in order to proceed with the storage process. Accordingly, the thermal energy requirement per mass CO₂ is given by:

$$Q_{\text{thermal}} = \frac{C_p m_{\text{sorbent}} \Delta T}{m_{\text{CO}_2}} + \frac{\Delta h_{\text{CO}_2} \Delta \sigma_{\text{CO}_2} + \Delta h_{\text{N}_2} \Delta \sigma_{\text{N}_2}}{m_{\text{CO}_2}}, \quad (1)$$

with the first term representing the sensible heat, which includes the parameters of the specific heat capacity C_p of the nanoporous material, the mass of the adsorbent m_{sorbent} , and the temperature difference between the adsorption and desorption process ΔT . The second term corresponds to the desorption heat, which is characterized by the working capacities $\Delta \sigma_i$ and the heat of adsorption Δh_i for both components, respectively. Pure component CO₂ and N₂ isotherms are obtained by fitting Langmuir isotherms to experimental data, if available, or data from molecular simulations. Moreover, Ideal Adsorbed Solution Theory (IAST),⁴⁵ is used to predict the behavior of a CO₂/N₂ gas mixture.

The second contribution to the parasitic energy is related to the compression of the desorbed CO₂ to either 1 bar, for utilization, or to 150 bar, for transportation and storage. The compression energy is mainly determined by the CO₂ purity and the desorption pressure. The pumping work, W_{comp} , is estimated by means of the NIST REFPROP database⁴⁶ and a linear regression based on this data. The isentropic efficiency of the pump is assumed to be 85% for gas below the supercritical point and 90% above it.



For the overall parasitic energy, we assume that heat is delivered by steam from the power plant and that the compressors are driven by the produced electricity directly. To calculate the total parasitic energy, the loss in the power plant's production, the compression work can be used directly, whereas for the heat, a typical turbine efficiency of 75%⁴⁷ and the Carnot efficiency η to convert thermal energy into electrical work have to be taken into account to translate the heat loss in an output loss:

$$E_{\text{parasitic}} = (0.75Q_{\text{thermal}}\eta_{\text{Carnot}}) + W_{\text{comp}} \quad (2)$$

This total energy requirement can be minimized by varying the final operation conditions of the desorption process, *i.e.* temperature and/or pressure. These desorption processes are referred to as Temperature–Pressure Swing Adsorption (TPSA), Temperature Swing Adsorption (TSA) or Pressure Swing Adsorption (PSA), depending which parameters are varied or fixed, in each case. TPSA operates at changing temperature and pressure, whereas TSA is characterized by constant pressure (set to $p_{\text{des}} = 1$ atm; this equals adsorption pressure) and PSA features a fixed temperature ($T_{\text{des}} = 333$ K). The considered desorption conditions range from $0.01 \text{ atm} < p_{\text{des}} < 3 \text{ atm}$ and $333 \text{ K} < T_{\text{des}} < 473 \text{ K}$ for pressure and temperature, respectively.

The database includes five classes of nanoporous materials, which are partially of hypothetical nature. These comprise metal–organic frameworks (MOFs),^{48–61} zeolitic imidazolate frameworks (ZIFs),^{62–64} porous polymer networks (PPNs),⁶⁵ zeolites²⁰ and cation exchanged zeolites (CEZs).⁴⁸ Moreover, two classes of hypothetical materials are considered: HMOFs – analogs of the well-studied MOF-5 (ref. 66) – and hypothetical ZIF structures. The class of CEZs is represented by Linde type A and type X zeolites where the Na-cations are partially exchanged by the alkaline earth metals magnesium and calcium. Alongside the fully coordinated MOFs, UMCM-1 and MOF-177, we also included MOFs with open metal sites, like the series M-MOF-74, CuBTC, and CuBTTri.^{11,26,67–69} Of special interest in recent years are also porous polymer networks (PPNs) tethered with different polyamines.^{70,71}

A full description of the methodology, the assumptions and the considered materials can be found in ref. 28.

4 Results & discussion

4.1 Minimum energy

Before we dig deeper into the performance of the different materials, it is worth looking at this problem from an elementary thermodynamical viewpoint.^{72,73} It is possible to calculate the minimum work to separate CO₂ from a two-component gas mixture, $W_{\text{min}}^{\text{sep}}$, using the molar entropy of an ideal gas mixture containing a mole fraction x of CO₂, according to

$$s^{\text{im}}(x) = -R[x \ln x + (1 - x)\ln(1 - x)] \quad (3)$$

$$-W_{\text{min}}^{\text{sep}} = -T_{\text{sep}}[n_{\text{em}}s^{\text{im}}(x_{\text{em}}) + n_{\text{cap}}s^{\text{im}}(x_{\text{cap}}) - n_{\text{flue}}s^{\text{im}}(x_{\text{flue}})], \quad (4)$$

where n_{em} , n_{cap} and n_{flue} are the number of moles of the emitted, captured and flue gas, respectively. A more detailed description of the calculation is given in the



ESI.† Fig. 2 shows the minimum separation energy landscape for a range of initial compositions (x_{flue} on the y-axis) and final compositions (x_{cap} on the x-axis) at 313 K. Note that the temperature of the flue gases is set to 313 K, whereas the ambient air is assumed to be 288 K. In this work, we focus on three sources of CO₂: coal-fired power plants, with 14% CO₂ in their flue gases, natural gas-fired power plants, emitting 4% CO₂, and ambient air, with some 400 ppm of CO₂ (0.04%). Moreover, we will consider three final purity requirements, 99%, 70% and 50% CO₂. These distinct compositions are shown in Fig. 2 on top of the minimum separation energy landscape. Moreover, Table 2 shows numerical values for the minimum energy requirement across the compositions we will investigate in this manuscript.

Direct air capture is situated in a very unfavorable region of this energy landscape and relative improvements are relatively small. For coal on the other hand, almost 50% minimum energy can be saved by lowering the purity of the final mixture from 99% to 50%. The energy requirement for upgrading the natural gas flue gas from 4% to 50% is almost identical to upgrading coal flue gas from 14% to 99%. This observation already opens perspectives to lower the cost of carbon capture from natural-gas fired power plants, given that applications for CO₂ at 50% purity can be found.

4.2 Influence of the purity requirement

Now, let us return to the selection of 62 nanoporous materials to consider carbon capture from the flue gas of coal-fired power plants. This flue gas has the highest CO₂ content and is therefore the most practical CO₂ source for post-combustion

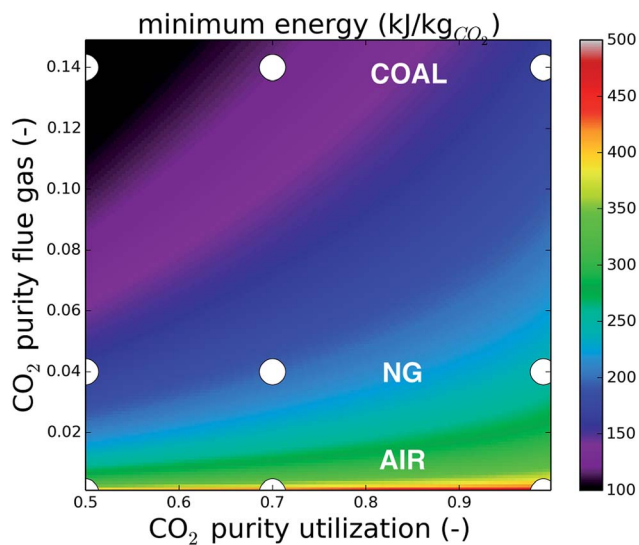


Fig. 2 The minimum energy (in kJ kg_{CO₂}⁻¹) to separate CO₂ from a binary mixture with an initial content (y-axis) to a final composition (x-axis), at 313 K and 1 bar. The points relevant in this manuscript are shown in white. The adsorption temperature in the case of direct air capture is 288 K in this manuscript, which lowers the minimum energy compared to the values in this figure.



Table 2 The minimum energy (in $\text{kJ kg}_{\text{CO}_2}^{-1}$) to separate CO_2 from a carbon source into a final composition. For coal (COAL) and natural gas (NG) flue gas, the temperature is set at 313 K, for direct air capture (AIR) at 288 K. The final pressure is set to 1 bar

CO_2 source		Final composition		
		99%	70%	50%
COAL	14%	168	119	89
NG	4%	245	197	166
AIR	0.04%	477	432	404

carbon capture. Fig. 3 shows the parasitic energy to capture CO_2 from coal flue gas as a function of the CO_2 Henry coefficient ($k_{\text{H,CO}_2}$) at 300 K, for the original requirement of CCS, with CO_2 at 99% purity and compressed to 150 bar and for the cases of 99%, 70% and 50% purity at 1 bar. First of all, the compression work from 1 to 150 bar comprises a large part of the parasitic energy and is hence an important part of the cost for CCS. Fig. S1 in the ESI† shows that this contribution to the parasitic energy exceeds 55%, especially for the best performing materials. Furthermore, lowering the purity requirement from 99% down to 70 or 50% further reduces the parasitic energy, although the relative improvements are much smaller than the parasitic energy reduction when omitting the 1 to 150 bar compression work.

The parasitic energy of the amine technology with monoethanolamine (MEA), $1065 \text{ kJ kg}_{\text{CO}_2}^{-1}$, is indicated with a solid line. Please note that for the MEA technology, the $353 \text{ kJ kg}_{\text{CO}_2}^{-1}$ compression work from 1 bar to 150 bar can be avoided, leaving the heating energy as the only contribution to the parasitic

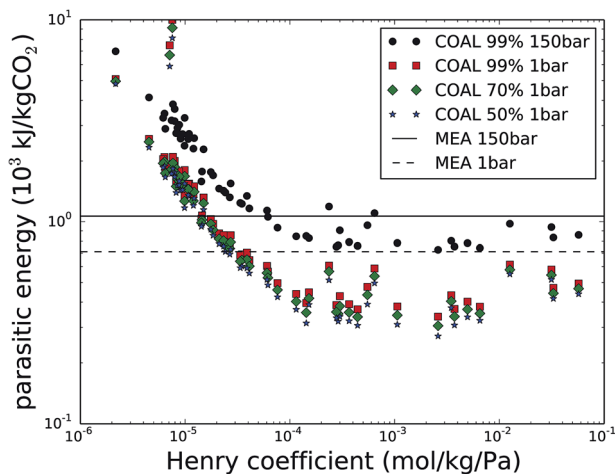


Fig. 3 The parasitic energy to capture CO_2 from coal flue gas as a function of the CO_2 Henry coefficient ($k_{\text{H,CO}_2}$) at 300 K. The original parasitic energy is given in black, with CO_2 at 99% purity and compressed to 150 bar. The cases for 99%, 70% and 50% purity at 1 bar are shown in red, green and blue respectively. Current MEA technology is marked as a solid line whereas the heating requirement of the MEA technology (excluding the compression from 1 to 150 bar) is shown as the dashed line.



energy (indicated with a dashed line in Fig. 3). This corresponds to $712 \text{ kJ kg}_{\text{CO}_2}^{-1}$, 33% lower than the original case. The purity of the final mixture however cannot be altered for MEA, as only CO_2 reacts with the amines and is therefore recovered at 100%. Within the scope of lowering the energy requirement of carbon capture for certain cases, of carbon utilization, solid adsorbents therefore provide larger margins for energy improvements than amines.

Table 3 shows the parasitic energy of the best performing materials for carbon capture from coal-fired power plants, for different final requirements. Mg-MOF-74 is the best candidate regardless of the final purity requirement. Not compressing the CO_2 to 150 bar saves $388 \text{ kJ kg}_{\text{CO}_2}^{-1}$ and lowers the parasitic energy with 54%. The net compression work is slightly higher than for MEA, as the CO_2 is not 100% pure when using Mg-MOF-74, so some N_2 has to be compressed as well. Further lowering the purity requirement decreases the parasitic energy from 339 to $271 \text{ kJ kg}_{\text{CO}_2}^{-1}$, an additional improvement of 20%. This is however much lower than the contribution of the 1–150 bar compression work.

The order of the other top-performing materials shifts slightly, depending on the imposed final conditions. The polyamine-tethered porous polymer network PPN-6- CH_2 -TETA⁷⁴ disappears from the selection at 70% and 50%. The same is true for the amine-functionalized metal-organic framework mmen-CuBTTri. Finally, for all considered cation-exchanged X-type zeolite (Mg, Na, Ca), the parasitic energy is very similar across the considered cases. The cation-exchanged zeolites are of high interest, as they combine their good carbon capture properties with commercial availability on the tonne scale. This is especially true for zeolite NaX, which is often used as a reference material for solid-adsorbent carbon capture.^{31,75–77} Finally, as the difference in carbon capture performance between the readily available NaX, and the yet-to-be commercialized Mg-MOF-74 is small, there is no reason to wait for the industrial production of novel materials in order to do carbon capture with solid adsorbents.

4.3 Influence of the carbon source

Extending the analysis to other sources of CO_2 , such as the flue gas of natural gas-fired power plants (with typically 4% CO_2 content) and ambient air (with CO_2 levels of 400 ppm), yields Fig. 4a. The 1–150 bar compression work is omitted in all cases and focus is on the imposed final purity. Fig. 4b shows the relative improvement between the parasitic energy for 50% *versus* 99% imposed purity for coal flue gases, natural gas flue gases and ambient air.

Table 3 Parasitic energy ($\text{kJ kg}_{\text{CO}_2}^{-1}$) of the five best performing materials for carbon capture from the flue gas of coal-fired power plants (COAL), for different final conditions

COAL							
150 bar, 99%	1 bar, 99%		1 bar, 70%		1 bar, 50%		
Mg-MOF-74	727	Mg-MOF-74	339	Mg-MOF-74	305	Mg-MOF-74	271
PPN-6- CH_2 -TETA	742	MgX	369	MgX	338	MgX	306
mmen-CuBTTri	752	NaX	370	NaX	340	NaX	308
NaX	754	PPN-6- CH_2 -TETA	380	CaX	344	CaX	310
MgX	760	CaX	381	NaA	350	Zn-MOF-74	315



As expected from the minimum energy calculations in Table 2, the parasitic energy for direct air capture is higher than that for carbon capture from natural gas flue gases or coal flue gases. Moreover, throughout the carbon sources, lower final purity requirements decrease the parasitic energy. The largest gains are found for direct air capture, for which a lot of the poorly performing materials exhibit relative improvements of 20–60% between 99% and 50% purity. This observation does not correspond to the conclusions from the minimum energy calculations, which predicted much lower improvements for direct air capture. Therefore, these large improvements are likely related to improvements in the process, rather than in the thermodynamics of air capture. For coal and natural gas, 10–20% energy can typically be saved, especially for the best performing materials.

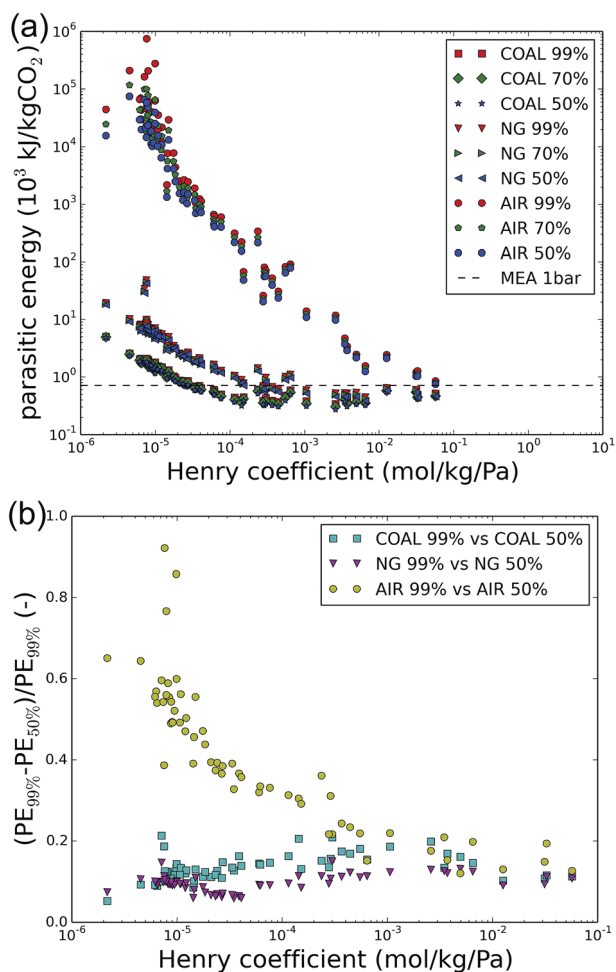


Fig. 4 (a) The parasitic energy to capture CO₂ from a coal-fired power plant (COAL), a natural gas-fired power plant (NG) or directly from air (AIR). Final purity requirements of 50, 70 and 99% are shown and 1 bar is maintained across all cases. (b) The relative gain in parasitic energy (PE) between the parasitic energy for 50% versus 99% imposed purity for coal flue gases, natural gas flue gases and ambient air.



Table 4 lists the five best performing materials for carbon capture from the flue gas of natural gas-fired power plants and directly from air, for the different final requirements. For natural gas flue gases, the functionalized PPN-6 materials have the lowest parasitic energy requirements. The top performer is polyamine-tethered porous polymer network PPN-6-CH₂-TETA, which was already mentioned among the five best candidates for carbon capture from coal flue gases. The metal-organic material (MOM) zinc hexafluorosilicate (SiF₆⁻²), SIFSIX-3-Zn, is highly selective for CO₂.^{78,79} In cases of lower purity, Mg-MOF-74 again acts as a promising candidate. And finally, the commercially available zeolite NaX is present in the top-5 materials for all purities. As opposed to the case of coal flue gases however, its performance is drastically lower than its competitors. For CO₂ capture at 99% purity and 150 bar, NaX has a parasitic energy of 925 kJ kg_{CO₂}⁻¹ whereas for PPN-6-CH₂-TETA, this is only 807 kJ kg_{CO₂}⁻¹.

For CO₂ capture directly from air, again the functionalized PPN-6 materials are suitable candidates, with PPN-6-CH₂-DETA as the material with the lowest parasitic energy. The differences in parasitic energy, even among the five best performers, are much larger than in the case for capture from coal or natural gas flue gases. Due to this wider spacing, the relative order of the materials does not change with changing purity requirements. Copper hexafluorosilicate (SiF₆⁻²), SIFSIX-3-Cu, is the second best material. And the zeolitic imidazolate framework ZIF-36-CAG also appears in this list. NaX is in the seventh place, with a parasitic energy of almost four times that of PPN-6-CH₂-DETA (*e.g.* 2889 kJ kg_{CO₂}⁻¹ for NaX *versus* 746 kJ kg_{CO₂}⁻¹ for PPN-6-CH₂-DETA in the case of 50% pure CO₂).

In Table 5, the parasitic energies of the best performing materials for each source of CO₂ are compared across the imposed purities. From Tables 3 and 4, we concluded that for CO₂ capture from coal-fired power plants, Mg-MOF-74 had the optimal properties, for natural gas-fired power plants, this was PPN-6-CH₂-TETA and for direct air capture PPN-6-CH₂-DETA. Lowering the final purity requirement

Table 4 Parasitic energy (kJ kg_{CO₂}⁻¹) of the five best performing materials for carbon capture from the flue gas of natural gas-fired power plants (NG) and directly from air (AIR), for different final conditions

150 bar, 99%	1 bar, 99%	1 bar, 70%	1 bar, 50%
NG			
PPN6-CH ₂ -TETA 807	PPN6-CH ₂ -TETA 444	PPN6-CH ₂ -TETA 416	PPN6-CH ₂ -TETA 389
PPN6-CH ₂ -TAEA 858	PPN6-CH ₂ -TAEA 494	PPN6-CH ₂ -TAEA 466	PPN6-CH ₂ -TAEA 439
PPN6-CH ₂ -DETA 880	PPN6-CH ₂ -DETA 515	Mg-MOF-74 486	Mg-MOF-74 454
SIFSIX-3-Zn 907	NaX 518	PPN6-CH ₂ -DETA 487	SIFSIX-3-Zn 455
NaX 925	SIFSIX-3-Zn 520	NaX 488	NaX 455
AIR			
PPN6-CH ₂ -DETA 1215	PPN6-CH ₂ -DETA 854	PPN6-CH ₂ -DETA 797	PPN6-CH ₂ -DETA 746
SIFSIX-3-Cu 1617	SIFSIX-3-Cu 1206	SIFSIX-3-Cu 1112	SIFSIX-3-Cu 1026
PPN6-CH ₂ -TAEA 1645	PPN6-CH ₂ -TAEA 1281	PPN6-CH ₂ -TAEA 1145	PPN6-CH ₂ -TAEA 1033
PPN6-CH ₂ -TETA 1948	PPN6-CH ₂ -TETA 1564	PPN6-CH ₂ -TETA 1398	PPN6-CH ₂ -TETA 1254
ZIF-36-CAG 3240	ZIF-36-CAG 2439	ZIF-36-CAG 2248	ZIF-36-CAG 2122



from 99% to 50% lowers the parasitic energy with almost 20% for Mg-MOF-74, whereas for PPN-6-CH₂-TETA and PPN-6-CH₂-DETA, the possible improvement is only 12%. Interestingly, for direct air capture, the parasitic energy requirement when purifying CO₂ from 400 ppm to 50% (746 kJ kg_{CO₂}⁻¹) is only marginally higher than the original parasitic energy requirement for CO₂ at 150 bar and 99% purity from coal flue gas (727 kJ kg_{CO₂}⁻¹). This indicates that direct air capture could be an attractive method to win CO₂ for specific cases of carbon utilization.

4.4 Importance of the final purity

In the previous sections, the CO₂ purity was a key factor in the discussion. The initial purity of CO₂ is determined by the composition of the flue gas or the air, and is therefore fixed by the carbon source. For the final purity of CO₂, we assumed that this is imposed by the operation of the carbon capture process. However, Fig. 5 reveals that this is not necessarily the case. In this plot, the results from Fig. 3 are color coded based on the final purity, after optimization of the parasitic energy. For a coal-fired power plant flue gas and with final purities set to 99%, 70% and 50%, only the best performing materials actually attain purities higher than 90%. Moreover, even when the purity requirement is gradually lowered (circles, squares and triangles at the same Henry coefficient), this barely influences the final purity. This indicates that the final purity of the recovered gas mixture is in fact imposed by the material properties (e.g. Henry coefficient), rather than by the operating conditions.

Fig. 6 shows the final purity as a function of the CO₂/N₂ selectivity of the material at adsorption conditions. A correlation exists between the CO₂ selectivity, a material property, and the final CO₂ purity, an outcome of the carbon capture process optimization. For coal flue gases, there is no significant difference between the different cases of imposed purity, which is a process parameter. Fig. S2 and S3 in the ESI† show that the points for natural gas flue gases and direct air capture are slightly further apart. In any case, the CO₂/N₂ selectivity of the used material influences the final purity more than the imposed purity does.

Moreover, a couple of interesting correlations with the purity are shown in Fig. 7. For clarity, the points are limited to the flue gases of coal-fired power plants. In the ESI,† the correlations are also given for natural gas flue gases and air. In Fig. 7a, the parasitic energy is plotted as a function of the final purity. Generally, the parasitic energy of a material decreases with increasing final purity. At very high final purity (>95%) however, the parasitic energies increase again, suggesting that the last percents are the most difficult to obtain. As was already clear from previous plots, the imposed purity also has an influence on the

Table 5 The parasitic energies of the best performing materials for each source of CO₂ are compared across the imposed final pressures and purities

	150 bar, 99%	1 bar, 99%	1 bar, 70%	1 bar, 50%
COAL	727	339	305	271
NG	807	444	416	389
AIR	1215	854	797	746



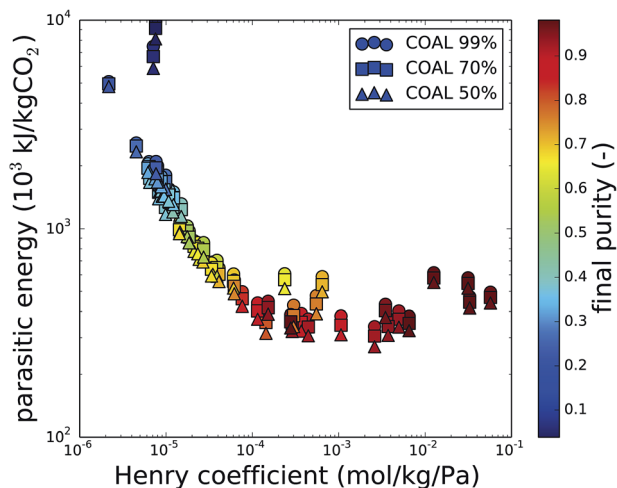


Fig. 5 The parasitic energy to capture CO_2 from coal flue gas as a function of the CO_2 Henry coefficient (k_{H,CO_2}) at 300 K, for CO_2 at 99%, 70% or 50% and 1 bar. The color code indicates the purity of the final gas mixture and shows that the purity requirement of 99% is only obtained for the best performing materials.

parasitic energy. For air capture (See Fig. S4 in the ESI[†]), most materials fail to meet the purity requirements, and hence result in parasitic energies of up to 4 orders of magnitude higher. Fig. 7b reveals an even clearer correlation between the final purity and the compression energy. Also here, both the imposed and the final purity have an effect on the compression energy. Therefore, it pays off to use a highly selective CO_2 adsorbent and lower the purity requirements later, rather

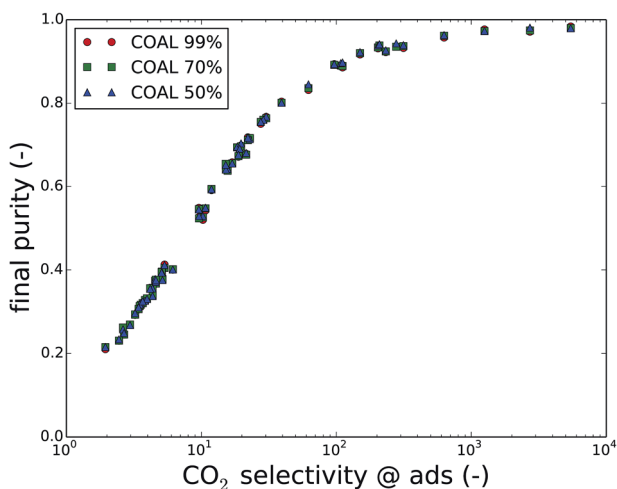


Fig. 6 The final purity as a function of the CO_2/N_2 selectivity of the material at adsorption conditions for carbon capture from coal flue gases at different imposed purities (99%, 70%, 50%). Note that the selectivity is a material property, the imposed purity an adaptable process parameter and the final purity an outcome of the process optimization.



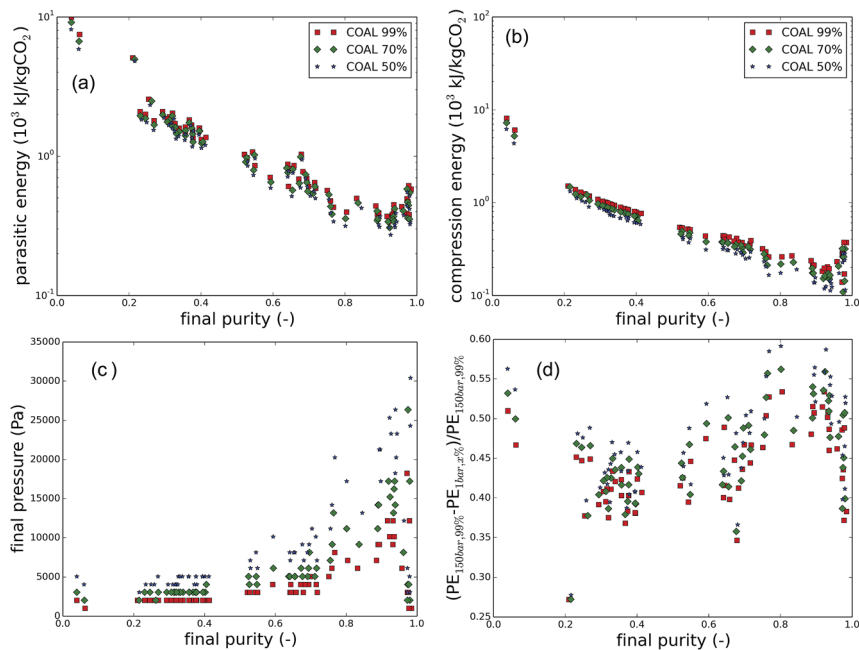


Fig. 7 (a) Parasitic energy as a function of the final purity, (b) the compression part of the parasitic energy as a function of the final purity, (c) the final pressure as a function of the final purity, (d) improvement of the parasitic energy, compared to the reference case of CO₂ at 150 bar and 99% imposed purity, all as a function of the final purity.

than using a less selective material CO₂ which would meet the goals. This is due to the compression work of the additional N₂ in the capture gas stream, as can be explained by Fig. 7c. The pressure at which CO₂ is recovered varies the most at high purity, dropping significantly across the imposed final purity, whereas at low purity, there is almost no improvement possible. Finally, in Fig. 7d, the parasitic energy across the different cases is compared with the original parasitic energy requirement for CO₂ at 99% and 150 bar shows that improvements of almost 60% are possible, especially when using coal flue gases as the carbon source. For direct air capture, the possible improvement of most materials is very low, but there are also some promising outliers at high purity (see Fig. S4 in ESI†).

These plots demonstrate that on the one hand, for one specific material, the parasitic energy is reduced when lower purity requirements are imposed. On the other hand, across the different materials, the trend is that the higher the purity of the final mixture, the lower the parasitic energy requirement. Although the purity is mainly imposed by the CO₂/N₂ selectivity of the material, which is fixed, there is some margin to tune the desorption pressure in order to lower the compression work of the regeneration. Therefore, the best strategy is to choose a highly selective material, which will hence yield a high final purity, ensuring a low parasitic energy among competing materials. Subsequently, the carbon capture process can be operated at a lower imposed purity, thereby reducing the parasitic energy.



Finally, when the desired purity is not met, a second adsorption step can be introduced. This step will be less energy-demanding, but will increase the parasitic energy of the overall process. This is an interesting extension for future work.

4.5 Case study: zeolite NaX

Zeolites are the work horses of the petrochemical industry, with widespread uses in catalysis and adsorption. Moreover, they have excellent regenerability and high stability in a wide range of environments. The commercial availability of zeolite NaX (or 13X) in particular, is truly unique among the considered materials.⁸⁰ Only a handful of MOFs have been commercialized so far, with prices that are many times that of zeolites.⁸¹ The other materials discussed in this manuscript are even further away from industrial applications.

Fig. 8 shows the Langmuir isotherms of CO₂ and N₂ in NaX, for carbon capture from coal flue gases (a) and directly from air (b) respectively. The numerical values for the desorption pressure, temperature and purity, as well as the compression, heating and parasitic energy are given in Table 6. For coal flue gases, the CO₂ loading during adsorption conditions is high among its competitors: almost 150 g_{CO₂} kg_{NaX}⁻¹, as is its selectivity (CO₂/N₂) with 314, resulting in a relatively low N₂ uptake. Optimizing the parasitic energy includes a trade-off between pulling a stronger vacuum, and hence evacuating more CO₂ from the material in one step, and keeping a higher desorption pressure, but requiring more energy for heating the material. The result of this optimization is the ideal amount of CO₂ recovered in one step, the difference in the CO₂ loading at adsorption and the CO₂ loading at desorption. This amount remains more or less constant across the imposed final purities. Lowering the final purity requirement means that the desorption sweep stream contains less CO₂ and that therefore, the driving force for CO₂ to leave the framework increases. As a result, the desorption isotherm moves towards higher total pressures, and hence, lower pressures are needed to evacuate the same amount of CO₂.

For direct air capture, Fig. 8b, the performance of NaX is suboptimal. In this case, the desorption pressure reached the lower limit of 1013 Pa for all cases. Therefore, the temperature has to be increased in order to recover enough CO₂. However, the CO₂ working capacity is still poor (only 20% of that for coal flue gases) and the final purity is below any of the imposed purities (around 34%). It is however possible to imagine carbon utilization at 30%, for instance to improve crop growth in greenhouses.^{35,36}

5 Cost from a sequestration and utilization viewpoint

As for any other chemical compound, there is a market for CO₂. As opposed to any other market however, the CO₂ supply however exceeds its demand with several orders of magnitude, as illustrated in Table 7. On the supply side, the 36 billion metric tons of CO₂ emitted in 2015. On the demand side, some 80 million metric tons of CO₂, most of which were used for Enhanced Oil Recovery (EOR). Currently, the CO₂ used for EOR is delivered from natural CO₂ reservoirs, rather than from carbon capture in power plants.⁸² Moreover, the recovered oil will in turn be converted into CO₂ emissions, giving rise to a net increase of CO₂ emissions to the



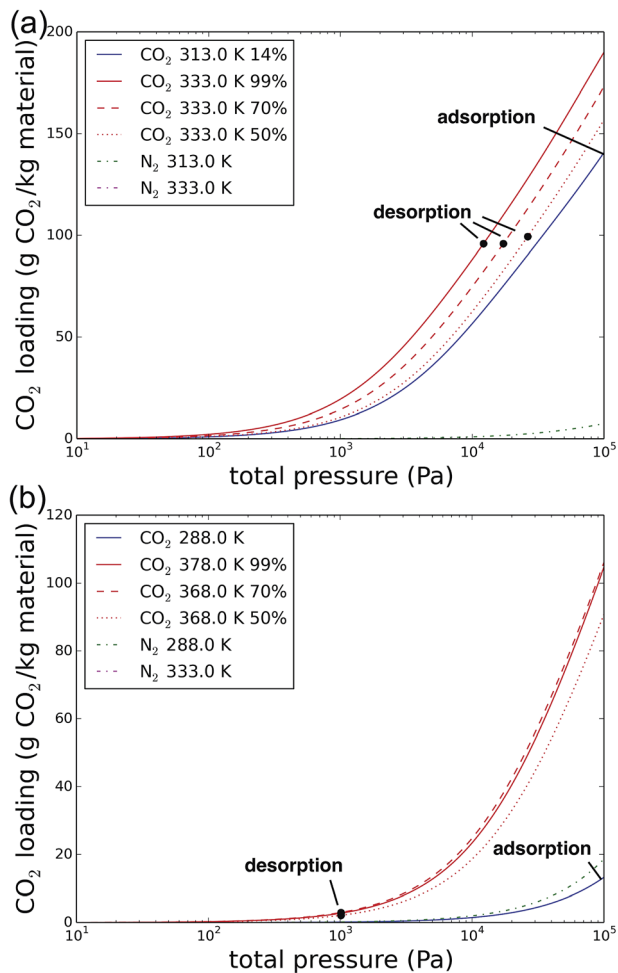


Fig. 8 The CO₂ and N₂ isotherms of NaX for carbon capture from coal-fired power plant flue gases (a) and direct air capture (b), respectively. The adsorption and different desorption conditions indicated are depending on the imposed purity of the CO₂.

atmosphere. Carbon capture and utilization for EOR will only be economically viable if the price of the carbon capture technology is lower than the price to extract CO₂ from natural reservoirs.

Market mechanisms alone will therefore not have a significant impact in order to lower CO₂ emissions to the atmosphere. That is why a growing number of countries, regions and cities are putting a price on CO₂ emissions, either through taxes or cap-and-trade mechanisms. Taxes straightforwardly provide incentives for polluters to reduce emissions, as it requires them to pay a fixed amount for each ton of CO₂ emitted. Cap-and-trade mechanisms require polluters to buy units of allowed emissions, which can be traded on the market, like any other commodity. Last year, this global market traded 7 billion metric tons of CO₂, a significant portion of the total emissions. At first sight, the total market for CCS



Table 6 Detailed results of the parasitic energy optimization for zeolite NaX for different initial CO₂ compositions and imposed final CO₂ purities. From left to right, the final pressure, temperature and purity (fin *P*, fin *T* and fin PUR), the compression energy from 1 to 150 bar (CE_{1-150 bar}), the parasitic energy (PE), the compression energy to 1 bar (CE), the heating energy (HE), the ratio of the heating energy and the parasitic energy (HE/PE) and the amount of CO₂ produced (CO₂ prod)

NaX	Fin <i>P</i> (Pa)	Fin <i>T</i> (K)	Fin PUR (-)	CE _{1-150 bar} (kJ kg _{CO₂} ⁻¹)	PE (kJ kg _{CO₂} ⁻¹)	CE (kJ kg _{CO₂} ⁻¹)	HE (kJ kg _{CO₂} ⁻¹)	CE/PE (-)	CO ₂ prod (kg m _{material} ⁻³)
COAL 99%	12 159	333	0.932	383	370	180	191	0.49	42.0
COAL 70%	17 225	333	0.937	381	340	149	190	0.44	42.1
COAL 50%	26 345	333	0.939	380	308	113	194	0.37	38.9
NG 99%	3040	333	0.886	406	518	313	205	0.60	42.7
NG 70%	4053	333	0.891	404	488	286	203	0.59	44.4
NG 50%	6080	333	0.889	405	455	250	205	0.55	42.5
AIR 99%	1013	378	0.338	1158	3412	1102	2309	0.32	8.7
AIR 70%	1013	368	0.333	1176	3103	1118	1984	0.36	8.5
AIR 50%	1013	368	0.353	1106	2889	1054	1835	0.36	9.2



Table 7 Global CO₂ emissions, traded emissions and CO₂ demand, in million metric tons per year

	Mton per year	Year	Reference
Global CO ₂ emissions	36 000	2015	83
Traded CO ₂ emissions	7000	2015	83
Demand for bulk CO ₂	80	2011	82
of which EOR	50	2011	82

is two orders of magnitude larger than the market for CCU. However, it only pays off to capture and sequester CO₂ emissions if the price to emit one ton of CO₂ is higher than the cost of a technology that avoids the emission of one ton of CO₂.

The trading prices of CO₂ in Europe and California were around 8 EUR and 12 USD per metric ton of CO₂ in 2015, whereas reported CCS costs are generally a few times higher (see Table 8). Moreover, on average, CO₂ emissions are priced at less than 10 USD per ton.⁸³ As a result, there is currently no market for CCS. With the current CO₂ trading prices, it is cheaper to dump CO₂ into the atmosphere and pay for it, than to avoid the emissions using CCS. In view of carbon utilization, it is not straightforward to find prices for bulk CO₂, as they are highly dependent on the source, and especially the distance from the source to the utilization site. We found a relatively low price for CO₂ from natural reservoirs, 15–19 USD per ton, but if new pipelines have to be build, this price will increase rapidly. Likewise, some chemical plants produce CO₂ with high purity and low costs (5–25 USD per ton), for instance from ammonia synthesis, but if the CO₂ is not used in the

Table 8 Various prices for CO₂

	ton CO ₂ ⁻¹	Currency	Year	Ref.
Carbon trading price				
EU	8	EUR	2015	84
California	12	USD	2015	85
CCS cost				
McKinsey	30–45	EUR	2008	2
BCG	45	EUR	2008	3
IEA (coal)	55	USD	2011	5
IEA (NG)	80	USD	2011	5
IPCC (coal)	9–44	USD	2005	86
IPCC (NG)	16–68	USD	2005	86
CO₂ production price				
Industrial grade	51	USD	2010	87
Natural reservoirs	15–19	USD	2011	82
Concentrated sources	20	USD	2011	82
Refinery	5–25	USD	2005	88
Consumer market	90–100	USD	2005	88
Social cost of carbon				
Interagency working group	43	USD	2013	89



proximity of the refinery, the cost will go up drastically. Moreover, when food-grade CO₂ is required, the price will also be higher than that of industrial grade, *etc.*

The central question in this manuscript is whether utilization at ambient pressure and lower purity further shift the economics in a more favorable direction. Since the parasitic energy is defined as the electricity loss of the power plant, the values in kJ kgCO₂⁻¹ can be converted to currency, by assuming an average electricity price. In 2015, the price for electricity for industrial customers was on average 0.12 EUR per kW per h in the European Union⁹⁰ and 0.07 USD in the United States of America.⁹¹ Note that there are large differences across these regions, and that no distinction is made between the relative costs of the different fuels (coal, natural gas, nuclear, renewable). Table 9 shows the monetary cost for the different carbon capture scenarios discussed above. These values do not include the capital cost, or the cost of transportation and sequestration, which is around 10 EUR per ton of CO₂.

The cost of 24 EUR per ton in Table 9 for carbon capture from the flue gas of coal-fired power plants is in agreement with the CCS costs reported in Table 8 when subtracting the transportation, storage and capital cost. The price of 14 USD is also relatively close to the 15–19 USD for CO₂ from natural CO₂ reservoirs, indicating that captured CO₂ from coal flue gases could possibly be used for EOR. Secondly, in Table 9, the difference between coal-fired power plants and natural gas-fired power plants is much smaller than what was expected from the policy reports in Table 8. For CO₂ capture directly from the air, we did not find reference prices in the literature, but the values in Table 9 are likely on the low side.

Most interestingly, not compressing the capture CO₂ from 1 bar to 150 bar saves around 13 EUR or 7 USD per ton of CO₂. Lowering the purity requirement from 99% to 70% or 50% can save 1–2 EUR or 1–2 USD, a modest improvement. The CO₂ price of 9 EUR or 5 USD per ton from coal flue gases is very interesting for applications where 1 bar CO₂ at 50% purity is desired. Finally, an interagency working group determined the US social cost of CO₂ at 43 USD per ton of CO₂ in 2020 assuming a 3% discount rate.^{89,92} At this socially desirable CO₂ price, sequestration and utilization of CO₂ would be economically viable.

Table 9 The minimal cost for carbon capture, assuming an average price of 0.07 USD per kW per h in the USA, or parasitic energies of the best performing materials for each source of CO₂, are compared across the imposed purities

	150 bar, 99%	1 bar, 99%	1 bar, 70%	1 bar, 50%
EUR				
COAL	24	11	10	9
NG	27	15	14	13
AIR	41	28	27	25
USD				
COAL	14	7	6	5
NG	16	9	8	8
AIR	24	17	15	14



6 Conclusions

Before CO₂ can be sequestered in underground reservoirs, it has to be purified to 99% and subsequently compressed to 150 bars, making up a large part of the cost for carbon capture and storage. For some examples of carbon capture and utilization however, compression to 150 bar is superfluous and the final purity requirement may be lower than 99%. The use of CO₂ to increase crop growth in greenhouses, algae growth for biofuel production or “Carbstone” are a few notable examples. In this manuscript, we investigate how far the parasitic energy – and therefore costs – of carbon capture can be lowered by considering less stringent final conditions. To find reliable trends, we screen a selection of over 60 materials, both synthesized and hypothetical, from a variety of material families.

First of all, we find that the compression work from 1 to 150 bar comprises up to 55% of the parasitic energy for CCS. Therefore, avoiding this compression work in a CCU application is by far the biggest possible improvement we identified. Moreover, lowering the purity requirement from 99% down to 50% further lowers the parasitic energy up to 20%. This stems from the fact that CO₂ can be recovered at a less strong vacuum, which requires less compression work.

When comparing different materials, two effects are at play. Across different materials, the higher the purity of the final mixture, the lower the parasitic energy requirement. For a single material however, the parasitic energy is lowered when lower purity requirements are imposed. There is a systematic distinction between the imposed purity and the final purity. We found that the CO₂/N₂ selectivity of the material, a fixed material property, and not the imposed final purity, one of the carbon capture process conditions, has the largest influence on the final CO₂ purity, a result of the parasitic energy minimization. When a lower purity is imposed, the final purity does not change significantly, but the desorption pressure is higher and the compression work of the regeneration is hence lower. Therefore, a highly selective material yields a high final purity, and the carbon capture process can subsequently be operated at a lower imposed purity, thereby lowering the parasitic energy.

As a result, the best performing materials remain the same across the considered final conditions, although the relative order may somewhat shift. For carbon capture from coal flue gases, natural gas flue gases and air, these materials are Mg-MOF-74, PPN-6-CH₂-TETA, and PPN-6-CH₂-DETA, respectively. Zeolite NaX combines its commercial availability with good performance, especially for coal flue gases and natural gas flue gases.

Interestingly, omitting the 1–150 bar compression and lowering the purity requirement reduces the parasitic energy of direct air capture almost to the parasitic energy to recover CO₂ at 99% purity and 150 bar from coal flue gases. Direct air capture could hence be attractive for specific cases of carbon utilization. Bear in mind that changing the purity requirement is not possible with the MEA technology, as 100% pure CO₂ is stripped from the amines in the regeneration step. This underlines the extra potential of solid adsorbent carbon capture in the case of CCU.

As a conclusion, when CO₂ is not sequestered at 150 bar and 99% purity, but instead utilized at ambient pressure and lower purities, the parasitic energy and associated cost of carbon capture may be reduced by almost 60%. This drastically



lowers the cost of carbon capture for utilization purposes. Moreover, the calculated price of CO₂ capture from coal flue gases, for utilization at 1 bar and 50% purity (9 EUR or 5 USD respectively), is competitive with CO₂ from natural reservoirs or refineries. This makes a strong case for carbon capture and utilization.

Acknowledgements

L. J. and V. V. S. acknowledge the Flemish Fund for Scientific Research (FWO12/ASP/151) and BELSPO in the frame of IAP/7/05 for financial support. V. V. S. acknowledges funding from the European Research Council under the European Community's Seventh Framework Programme FP7 (2007–2013) ERC grant agreement number 240483, of the European Union's Horizon 2020 research and innovative programme under consolidator ERC grant agreement number 647755-DYNPOR (2015–2020). B. S. was supported by the Center for Gas Separations Relevant to Clean Energy Technologies, an Energy Frontier Research Center funded by the U.S. Department of Energy, Office of Science, and Office of Basic Energy Sciences, under Award Number DE-SC0001015. J. M. H. was supported by the Deutsche Forschungsgemeinschaft (DFG, priority program SPP 1570).

References

- 1 A. S. Bhowan and B. C. Freeman, *Environ. Sci. Technol.*, 2011, **45**, 8624–8632.
- 2 McKinsey&Company, *Carbon Capture & Storage: Assessing the Economics*, 2008, <http://www.globalecsinstitute.com/publications/carbon-capture-storage-assessing-economics>.
- 3 Boston Consulting Group, *Carbon Capture and Storage: A Solution to the Problem of Carbon Emissions*, 2008, <https://www.bcg.com/documents/file15263.pdf>.
- 4 British Petroleum, *BP Energy Outlook 2035*, 2015, <https://www.bp.com/energyoutlook>.
- 5 International Energy Agency, *World Energy Outlook 2013 factsheet*, 2013, <http://www.worldenergyoutlook.org/resources/factsheets/>.
- 6 G. T. Rochelle, *Science*, 2009, **325**, 1652–1654.
- 7 G. Rochelle, E. Chen, S. Freeman, D. Van Wagener, Q. Xu and A. Voice, *Chem. Eng. J.*, 2011, **171**, 725–733.
- 8 D. M. D'Alessandro, B. Smit and J. R. Long, *Angew. Chem., Int. Ed.*, 2010, **49**, 6058–6082.
- 9 S. Choi, J. Drese and C. Jones, *ChemSusChem*, 2009, **2**, 796–854.
- 10 G. Sneddon, A. Greenaway and H. H. P. Yiu, *Adv. Energy Mater.*, 2014, **4**, 1301873.
- 11 A. O. Yazaydin, R. Q. Snurr, T.-H. Park, K. Koh, J. Liu, M. D. LeVan, A. I. Benin, P. Jakubczak, M. Lanuza, D. B. Galloway, J. J. Low and R. R. Willis, *J. Am. Chem. Soc.*, 2009, **131**, 18198–18199.
- 12 M. Fernandez, P. G. Boyd, T. D. Daff, M. Z. Aghaji and T. K. Woo, *J. Phys. Chem. Lett.*, 2014, **5**, 3056–3060.
- 13 N. T. T. Nguyen, H. Furukawa, F. Gándara, H. T. Nguyen, K. E. Cordova and O. M. Yaghi, *Angew. Chem., Int. Ed.*, 2014, **53**, 10645–10648.
- 14 B. Wang, A. P. Cote, H. Furukawa, M. O'Keeffe and O. M. Yaghi, *Nature*, 2008, **453**, 207–211.



- 15 D. Yuan, W. Lu, D. Zhao and H.-C. Zhou, *Adv. Mater.*, 2011, **23**, 3723–3725.
- 16 W. Lu, D. Yuan, D. Zhao, C. I. Schilling, O. Plietzsch, T. Muller, S. Bräse, J. Guenther, J. Blümel, R. Krishna, Z. Li and H.-C. Zhou, *Chem. Mater.*, 2010, **22**, 5964–5972.
- 17 W. Lu, W. M. Verdegaal, J. Yu, P. B. Balbuena, H.-K. Jeong and H.-C. Zhou, *Energy Environ. Sci.*, 2013, **6**, 3559.
- 18 T.-H. Bae, M. R. Hudson, J. A. Mason, W. L. Queen, J. J. Dutton, K. Sumida, K. J. Micklash, S. S. Kaye, C. M. Brown and J. R. Long, *Energy Environ. Sci.*, 2013, **6**, 128–138.
- 19 J. Kim, M. Abouelnasr, L.-C. Lin and B. Smit, *J. Am. Chem. Soc.*, 2013, **135**, 7545–7552.
- 20 M. W. Deem, R. Pophale, P. A. Cheeseman and D. J. Earl, *J. Phys. Chem. C*, 2009, **113**, 21353–21360.
- 21 M. M. F. Hasan, E. L. First and C. A. Floudas, *Phys. Chem. Chem. Phys.*, 2013, **15**, 17601–17618.
- 22 L. Joos, K. Lejaeghere, J. Huck, V. Van Speybroeck and B. Smit, *Energy Environ. Sci.*, 2015, **8**, 2480–2491.
- 23 H.-C. Zhou, J. R. Long and O. M. Yaghi, *Chem. Rev.*, 2012, **112**, 673–674.
- 24 V. Van Speybroeck, K. Hemelsoet, L. Joos, M. Waroquier, R. G. Bell and C. R. A. Catlow, *Chem. Soc. Rev.*, 2015, **44**, 7044–7111.
- 25 J. Li, A. Corma and J. Yu, *Chem. Soc. Rev.*, 2015, **44**, 7112–7127.
- 26 Y.-S. Bae and R. Q. Snurr, *Angew. Chem., Int. Ed.*, 2011, **50**, 11586–11596.
- 27 L.-C. Lin, A. H. Berger, R. L. Martin, J. Kim, J. A. Swisher, K. Jariwala, C. H. Rycroft, A. S. Bhowan, M. Deem, M. Haranczyk and B. Smit, *Nat. Mater.*, 2012, **11**, 633–641.
- 28 J. M. Huck, L.-C. Lin, A. H. Berger, M. N. Shahrak, R. L. Martin, A. S. Bhowan, M. Haranczyk, K. Reuter and B. Smit, *Energy Environ. Sci.*, 2014, **7**, 4132–4146.
- 29 C. Simon, R. Mercado, S. Schnell, B. Smit and M. Haranczyk, *Chem. Mater.*, 2015, **27**, 4459–4475.
- 30 C. Simon, J. Kim, D. Gomez-Gualdrón, J. Camp, Y. Chung, R. Martin, R. Mercado, M. Deem, D. Gunter, M. Haranczyk, D. Sholl, R. Snurr and B. Smit, *Energy Environ. Sci.*, 2015, **8**, 1190–1199.
- 31 L. Joos, J. A. Swisher and B. Smit, *Langmuir*, 2013, **29**, 15936–15942.
- 32 S. Heylen, L. Joos, T. N. Parac-Vogt, V. Van Speybroeck, C. E. A. Kirschhock and J. A. Martens, *Angew. Chem., Int. Ed.*, 2012, **51**, 11010–11013.
- 33 B. Smit, J. R. Reimer, C. M. Oldenburg and I. C. Bourg, *Introduction to Carbon Capture and Sequestration*, Imperial College Press, London, 1st edn, 2014.
- 34 S. A. Scott, M. P. Davey, J. S. Dennis, I. Horst, C. J. Howe, D. J. Lea-Smith and A. G. Smith, *Curr. Opin. Biotechnol.*, 2010, **21**, 277–286.
- 35 S. A. Prior, G. B. Runion, S. C. Marble, H. H. Rogers, C. H. Gilliam and H. A. Torbert, *HortScience*, 2011, **46**, 158–162.
- 36 H. H. Rogers, G. Runion and S. V. Krupa, *Environ. Pollut.*, 1994, **83**, 155–189.
- 37 C. Meyer, M. G. Wichmann and T. S. Spengler, *J. Econ. Bus.*, 2016, 1–36.
- 38 D. H. Abbott and C. E. Albright, *J. Laser Appl.*, 1994, **6**, 69.
- 39 W. Klemm and R. Berger, *Cem. Concr. Res.*, 1972, **2**, 567–576.
- 40 C. Wang, L. Lü and M. Xu, *J. Med. Entomol.*, 2012, **49**, 1076–1083.
- 41 E. N. T. George Douglas Hobson, *Introduction to Petroleum Geology*, Scientific Press Limited, 1st edn, 1975.
- 42 E. Davidson, *Process of extinguishing fires, US Pat.*, 1,010,870, 1911.



- 43 Q. Wang, J. Luo, Z. Zhong and A. Borgna, *Energy Environ. Sci.*, 2011, **4**, 42–55.
- 44 A. Kumar, D. G. Madden, M. Lusi, K.-J. Chen, E. A. Daniels, T. Curtin, J. J. Perry and M. J. Zaworotko, *Angew. Chem., Int. Ed.*, 2015, **54**, 14372–14377.
- 45 A. L. Myers and J. M. Prausnitz, *AIChE J.*, 1965, **11**, 121–127.
- 46 E. W. Lemmon, M. L. Huber and M. O. McLinden, *NIST Reference Fluid Thermodynamic and Transport Properties Database (REFPROP): Version 9.0*, 2010, <http://www.nist.gov/srd/nist23.cfm>.
- 47 S. A. Freeman, R. Dugas, D. H. Van Wagener, T. Nguyen and G. T. Rochelle, *Int. J. Greenhouse Gas Control*, 2010, **4**, 119–124.
- 48 D. M. D'Alessandro, B. Smit and J. R. Long, *Angew. Chem., Int. Ed.*, 2010, **49**, 6058–6082.
- 49 K. Sumida, D. L. Rogow, J. A. Mason, T. M. McDonald, E. D. Bloch, Z. R. Herm, T.-H. Bae and J. R. Long, *Chem. Rev.*, 2012, **112**, 724–781.
- 50 J.-R. Li, J. Sculley and H.-C. Zhou, *Chem. Rev.*, 2012, **112**, 869–932.
- 51 J. M. Simmons, H. Wu, W. Zhou and T. Yildirim, *Energy Environ. Sci.*, 2011, **4**, 2177.
- 52 J. A. Mason, K. Sumida, Z. R. Herm, R. Krishna and J. R. Long, *Energy Environ. Sci.*, 2011, **4**, 3030.
- 53 T. M. McDonald, D. M. D'Alessandro, R. Krishna and J. R. Long, *Chem. Sci.*, 2011, **2**, 2022.
- 54 P. Aprea, D. Caputo, N. Gargiulo, F. Iucolano and F. Pepe, *J. Chem. Eng. Data*, 2010, **55**, 3655–3661.
- 55 P. Chowdhury, S. Mekala, F. Dreisbach and S. Gumma, *Microporous Mesoporous Mater.*, 2012, **152**, 246–252.
- 56 P. D. C. Dietzel, R. Blom and H. Fjellvåg, *Eur. J. Inorg. Chem.*, 2008, **2008**, 3624–3632.
- 57 P. D. C. Dietzel, V. Besikiotis and R. Blom, *J. Mater. Chem.*, 2009, **19**, 7362.
- 58 L. Grajciar, A. D. Wiersum, P. L. Llewellyn, J.-S. Chang and P. Nachtigall, *J. Phys. Chem. C*, 2011, **115**, 17925–17933.
- 59 Z. Liang, M. Marshall and A. L. Chaffee, *Energy Fuels*, 2009, **23**, 2785–2789.
- 60 B. Mu, P. M. Schoenecker and K. S. Walton, *J. Phys. Chem. C*, 2010, **114**, 6464–6471.
- 61 Z. R. Herm, J. A. Swisher, B. Smit, R. Krishna and J. R. Long, *J. Am. Chem. Soc.*, 2011, **133**, 5664–5667.
- 62 R. Banerjee, H. Furukawa, D. Britt, C. Knobler, M. O'Keeffe and O. M. Yaghi, *J. Am. Chem. Soc.*, 2009, **131**, 3875–3877.
- 63 H. Amrouche, S. Aguado, J. Pérez-Pellitero, C. Chizallet, F. Siperstein, D. Farrusseng, N. Bats and C. Nieto-Draghi, *J. Phys. Chem. C*, 2011, **115**, 16425–16432.
- 64 H. Huang, W. Zhang, D. Liu, B. Liu, G. Chen and C. Zhong, *Chem. Eng. Sci.*, 2011, **66**, 6297–6305.
- 65 D. Yuan, W. Lu, D. Zhao and H.-C. Zhou, *Adv. Mater.*, 2011, **23**, 3723–3725.
- 66 R. L. Martin, L.-C. Lin, K. Jariwala, B. Smit and M. Haranczyk, *J. Phys. Chem. C*, 2013, **117**, 12159–12167.
- 67 P. Bai, M. Tsapatsis and J. I. Siepmann, *Langmuir*, 2012, **28**, 15566–15576.
- 68 J. W. Yoon, Y.-K. Seo, Y. K. Hwang, J.-S. Chang, H. Leclerc, S. Wuttke, P. Bazin, A. Vimont, M. Daturi, E. Bloch, P. L. Llewellyn, C. Serre, P. Horcajada, J.-M. Grenèche, A. E. Rodrigues and G. Férey, *Angew. Chem., Int. Ed.*, 2010, **49**, 5949–5952.



- 69 J. Liu, P. K. Thallapally, B. P. McGrail, D. R. Brown and J. Liu, *Chem. Soc. Rev.*, 2012, **41**, 2308–2322.
- 70 W. Lu, D. Yuan, J. Sculley, D. Zhao, R. Krishna and H.-C. Zhou, *J. Am. Chem. Soc.*, 2011, **133**, 18126–18129.
- 71 W. Lu, J. P. Sculley, D. Yuan, R. Krishna, Z. Wei and H.-C. Zhou, *Angew. Chem., Int. Ed.*, 2012, **51**, 7480–7484.
- 72 K. S. Lackner, *Energy*, 2013, **50**, 38–46.
- 73 Direct Air Capture of CO₂ with Chemicals: A Technology Assessment for the APS Panel on Public Affairs, 2011, <https://www.aps.org/policy/reports/assessments/upload/dac2011.pdf>.
- 74 W. Lu, J. P. Sculley, D. Yuan, R. Krishna, Z. Wei and H.-C. Zhou, *Angew. Chem., Int. Ed.*, 2012, **51**, 7480–7484.
- 75 J. Mérel, M. Clausse and F. Meunier, *Environ. Prog.*, 2006, **25**, 327–333.
- 76 D. Ko, R. Siriwardane and L. T. Biegler, *Ind. Eng. Chem. Res.*, 2003, **42**, 339–348.
- 77 K. T. Chue, J. N. Kim, Y. J. Yoo, S. H. Cho and R. T. Yang, *Ind. Eng. Chem. Res.*, 1995, **34**, 591–598.
- 78 P. Nugent, Y. Belmabkhout, S. D. Burd, A. J. Cairns, R. Luebke, K. Forrest, T. Pham, S. Ma, B. Space, L. Wojtas, M. Eddaoudi and M. J. Zaworotko, *Nature*, 2013, **495**, 80–84.
- 79 O. Shekhah, Y. Belmabkhout, Z. Chen, V. Guillermin, A. Cairns, K. Adil and M. Eddaoudi, *Nat. Commun.*, 2014, **5**, 4228.
- 80 Transparency Market Research, Global Synthetic Zeolite Market, <http://www.transparencymarketresearch.com/pressrelease/synthetic-zeolites-market.htm>.
- 81 B. Yilmaz, N. Trukhan and U. Müller, *Chin. J. Catal.*, 2012, **33**, 3–10.
- 82 Accelerating the Uptake of CCS: Industrial Use of Capture Carbon Dioxide, 2011, <https://www.hub.globalccsinstitute.com/publications/>.
- 83 State and Trends of Carbon Pricing, 2015, <http://www.worldbank.org/content/dam/Worldbank/document/Climate/State-and-Trend-Report-2015.pdf>.
- 84 European Emission Allowances—Global Environmental Exchange, 2015, <https://www.eex.com/en/market-data/emission-allowances/spot-market/european-emission-allowances>.
- 85 California Carbon Dashboard, 2015, <http://www.calcarbodash.org>.
- 86 *IPCC Special Report on Carbon Capture and Storage*, 2005.
- 87 Carbon Dioxide Enhanced Oil Recovery – Untapped Domestic Energy Supply and Long Term Carbon Storage Solution, 2010, https://www.netl.doe.gov/file/%20library/research/oil-gas/small_CO2_EOR_Primer.pdf.
- 88 *Carbon Dioxide Apps Are Key In Ethanol Project Developments*, 2011, <http://www.ethanolproducer.com/articles>.
- 89 Social Cost of Carbon for Regulatory Impact Analysis? Under Executive Order 12866, 2013, <http://www.1.usa.gov/18ftAsH>.
- 90 Energy price statistics, 2015, http://www.eceuropa.eu/eurostat/statistics-explained/index.php/Energy_price_statistics.
- 91 Electric Power Monthly, 2015, <http://www.eia.gov/electricity/monthly/pdf/epm.pdf>.
- 92 A. Fraas, R. Lutter, S. Dudley, T. Gayer, J. Graham, J. F. Shogren and W. K. Viscusi, *Science*, 2016, **351**, 569.

