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Comparison of group 4 and thorium M(IV) substituted cyclopentadienyl silanide complexes†

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We report the synthesis and characterisation of a series of M(IV) substituted cyclopentadienyl hypersilanide complexes of the general formula $[M(\text{Cp}^R)_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{X})]$ ($M = \text{Hf, Th}$; $\text{Cp}^R = \text{Cp}'$, $\{\text{C}_5\text{H}_4(\text{SiMe}_3)\}$ or Cp'' , $\{\text{C}_5\text{H}_3(\text{SiMe}_3)_{2-1,3}\}$; $\text{X} = \text{Cl, C}_3\text{H}_5$). The separate salt metathesis reactions of $[M(\text{Cp}^R)_2(\text{Cl})_2]$ ($M = \text{Zr or Hf}$, $\text{Cp}^R = \text{Cp}'$; $M = \text{Hf or Th}$, $\text{Cp}^R = \text{Cp}''$) with equimolar $\text{K}\{\text{Si}(\text{SiMe}_3)_3\}$ gave the respective mono-silanide complexes $[M(\text{Cp}^R)_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ ($M = \text{Zr, 1; Hf, 2}$), $[\text{Hf}(\text{Cp}'')(\text{Cp}')\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ (**3**) and $[\text{Th}(\text{Cp}'')_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ (**4**), with only a trace amount of **3** presumably formed via silatropic and sigmatropic shifts; the synthesis of **1** from $[\text{Zr}(\text{Cp}')_2(\text{Cl})_2]$ and $\text{Li}\{\text{Si}(\text{SiMe}_3)_3\}$ has been reported previously. The salt elimination reaction of **2** with one equivalent of allylmagnesium chloride gave $[\text{Hf}(\text{Cp}')_2\{\text{Si}(\text{SiMe}_3)_3\}(\eta^3\text{-C}_3\text{H}_5)]$ (**5**), whilst the corresponding reaction of **2** with equimolar benzyl potassium yielded $[\text{Hf}(\text{Cp}')_2(\text{CH}_2\text{Ph})_2]$ (**6**) together with a mixture of other products, with elimination of both KCl and $\text{K}\{\text{Si}(\text{SiMe}_3)_3\}$. Attempts to prepare isolated $[M(\text{Cp}^R)_2\{\text{Si}(\text{SiMe}_3)_3\}]^+$ cations from **4** or **5** by standard abstraction methodologies were unsuccessful. The reduction of **4** with KC_8 gave the known Th(III) complex, $[\text{Th}(\text{Cp}'')_3]$. Complexes **2–6** were characterised by single crystal XRD, whilst **2, 4** and **5** were additionally characterised by ^1H , $^{13}\text{C}\{^1\text{H}\}$ and $^{29}\text{Si}\{^1\text{H}\}$ NMR spectroscopy, ATR-IR spectroscopy and elemental analysis. In order to probe differences in M(IV)–Si bonds for d- and f-block metals we studied the electronic structures of **1–5** by density functional theory calculations, showing M–Si bonds of similar covalency for Zr(IV) and Hf(IV), and less covalent M–Si bonds for Th(IV).

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Introduction

Transition metal (TM) silicon chemistry is well-established,¹ with applications ranging from homogeneous (hydro)silylation catalysts^{2–5} to TM silicides in catalysis, ceramics and microelectronics.⁶ f-Block silicon chemistry is currently underdeveloped in comparison,^{1,7} but potential uses include lanthanide (Ln) silanide ($\{\{\text{SiR}_3\}\}$) complexes in olefin polymerisation catalysts,^{8,9} Ln silicides as additives in low-alloy steels,¹⁰ and high-density actinide (An) silicide nuclear fuels.^{11–14} Given that the physical and chemical properties of Ln and An complexes can differ markedly from each other and related early TM complexes,^{15,16} it follows that comparative studies of their electronic structures and the relative amount of covalency

in M–Si bonds are crucial for the development of novel applications.

Since the discovery nearly 40 years ago that group 4 TM bent metallocenes can promote the dehydrogenative polymerisation of silanes,¹⁷ M(IV) silanide complexes of the general formula $[M(\text{Cp}^R)_2\{\text{SiR}_3\}(\text{X})]$ ($M = \text{Ti, Zr, Hf}$; $\text{Cp}^R =$ substituted cyclopentadienyl, $\{\text{C}_5\text{R}_5\}$; $\text{X} =$ anion) have been studied extensively;¹ however, f-block analogues of these complexes have not previously been isolated. As the +4 oxidation state is limited for organometallic Ln complexes¹⁵ and given the increased radiological hazard and decreased stability of An(IV) complexes across the series,¹⁶ we targeted a Th(IV) silanide complex for comparison with group 4 M(IV) analogues. Whilst there are >130 structurally authenticated examples of Ti, Zr and Hf complexes containing one or more M–Si bonds, and >90 complexes containing Ln–Si bonds, for An silicon chemistry there are only 1 Th–Si and 6 U–Si bonds reported to date.¹⁸ Some of us recently reported the synthesis and characterisation of the sole thorium silanide complex, $[\text{Th}(\text{Cp}')_3\{\text{Si}(\text{SiMe}_3)_3\}]$ ($\text{Cp}' = \{\text{C}_5\text{H}_4(\text{SiMe}_3)\}$) and showed by a variety of calculated covalency metrics that the Th–Si bond was less covalent than the corresponding U–Si bond of the isostructural U(IV) complex $[\text{U}(\text{Cp}')_3\{\text{Si}(\text{SiMe}_3)_3\}]$.¹⁹ Our previously reported

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reactive towards KCH_2Ph than **2**, though we cannot discount ligand scrambling processes occurring. A small-scale reaction of **2** with benzyl potassium in C_6D_6 monitored by $^{29}\text{Si}\{^1\text{H}\}$ NMR spectroscopy exhibited a single resonance at -86 ppm after 15 min of reaction time, indicating that a product with a Hf–Si bond is present in the reaction mixture, but this could not be structurally authenticated; the product distribution of this small-scale reaction is likely not representative of the scaled-up version due to differences in solvent, substrate concentration and reaction time.

In efforts to prepare isolated $[\text{M}(\text{Cp}^R)_2\{\text{Si}(\text{SiMe}_3)_3\}]^+$ cations by adapting literature abstraction methodologies, complex **4** was treated with $[\text{CPh}_3][\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]^{29}$ and two equivalents of Et_3SiH (in order to form a silylium cation *in situ* that abstracts Cl^- to form strong Si–Cl bonds as an enthalpic driving force)³⁰ in benzene, and complex **5** was treated with $[\text{NEt}_3\text{H}][\text{BPh}_4]^{31}$ in either toluene or THF (in order to provide an entropic driving force by the formation of propene and trimethylamine *via* a protonolysis reaction). No product could be isolated from the former reaction, whilst no reaction was observed between **5** and the ammonium reagent in both solvent systems investigated. Finally, in an attempt to target the Th(III) complex $[\text{Th}(\text{Cp}^R)_2\{\text{Si}(\text{SiMe}_3)_3\}]$, which would be structurally analogous to previously reported Ti, Zr, La, Ce, Nd and U analogues,²⁰ complex **4** was reacted with 1.2 eq. of KC_8 ³² in DME. Following work-up and recrystallisation from pentane the previously reported homoleptic Th(III) complex $[\text{Th}(\text{Cp}^R)_3]^{33,34}$ was isolated in 59% crystalline yield, presumably by ligand scrambling of $[\text{Th}(\text{Cp}^R)_2(\text{X})]$ following elimination of KX ($\text{X} = \text{Cl}, \{\text{Si}(\text{SiMe}_3)_3\}$) in accord with the previously reported reduction of $[\text{Th}(\text{Cp}^R)_2(\text{Cl})_2]$ in THF over excess Na/K alloy.³²

Elemental analysis results obtained for **1**, **2**, **4** and **5** were in reasonable agreement with expected values, and their ATR-IR spectra (see ESI Fig. S39–S42†) exhibit absorption features that are similar to each other and those previously reported for the closely related M(III) complexes $[\text{M}(\text{Cp}^R)_2\{\text{Si}(\text{SiMe}_3)_3\}]$ ($\text{M} = \text{Ti}, \text{Zr}, \text{La}, \text{Ce}, \text{Nd}, \text{U}$).²⁰ The ATR-IR spectrum of **5** additionally contains the expected resonance at 1531 cm^{-1} for a Hf(IV) η^3 -allyl group, which is at a higher frequency to that previously reported for $[\text{Hf}(\text{Cp}^*)(\text{C}_4\text{H}_4\text{BN}^i\text{Pr}_2)(\eta^3\text{-C}_3\text{H}_5)]$ ($\tilde{\nu} = 1504\text{ cm}^{-1}$).³⁵

NMR spectroscopy

Complexes **1**, **2**, **4** and **5** were characterised by ^1H , $^{13}\text{C}\{^1\text{H}\}$ and $^{29}\text{Si}\{^1\text{H}\}$ NMR spectroscopy (see Table 1 and ESI Fig. S1–S38†). ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR chemical shifts were previously reported for **1**²¹ and the M(IV) precursors $[\text{M}(\text{Cp}^R)_2(\text{Cl})_2]$,^{20,22–24} but in order to consistently use C_6D_6 as the solvent we include these data for $[\text{M}(\text{Cp}^R)_2(\text{Cl})_2]$ herein, along with their $^{29}\text{Si}\{^1\text{H}\}$ NMR spectral data for completeness. The spectral assignments were confirmed by correlation experiments for **1**, **2**, **4** and **5**.

The ^1H NMR spectra of $[\text{M}(\text{Cp}^R)_2(\text{Cl})_2]$ exhibited three major signals in the expected ratio of 4:4:18, with typical second order AA'BB' four spin systems for the two Cp' ring proton environments and a singlet for the trimethylsilyl groups. The ^1H NMR spectra of **1**²¹ and **2** show six major signals in a ratio

of 2:2:2:2:27:18, with two sets of AA'BB' patterns for the Cp' ring protons; the inequivalent 2,5- and 3,4-Cp' ring proton resonances arise from the diminished symmetry in **1** and **2** compared with parent $[\text{M}(\text{Cp}^R)_2(\text{Cl})_2]$,²² which have mirror planes bisecting the metal and Cp' centroids. The ^1H NMR spectrum of **5** shows a similar pattern of signals to **1** and **2**, as well as the expected doublet (4H) signal for the methylene and quintet (1H) resonance for the methine of the allyl ligand. This AX₄ pattern is characteristic of fluxional behaviour, interconverting between η^3 - and η^1 -bound forms faster than the NMR timescale at this temperature; previously reported mono-allyl Hf(IV) complexes can show inequivalent *syn*- and *anti*-methylene groups in an AM₂X₂ pattern due to stronger η^3 -binding locking this conformation in solution, *e.g.* $[\text{Hf}(\text{Cp}^*)(\text{C}_4\text{H}_4\text{BN}^i\text{Pr}_2)(\eta^3\text{-C}_3\text{H}_5)]$ ($\text{Cp}^* = \text{C}_5\text{Me}_5$; $\delta_{\text{H}} = 6.99$ ppm, m, 1H, CH; 3.66 ppm, d, 2H, $^3J_{\text{HH}} = 15.4$ Hz, CH_2 -*anti*; 1.70 ppm, d, 2H, $^3J_{\text{HH}} = 9.0$ Hz, CH_2 -*syn*)³⁵ and $[\text{Hf}(\text{Cp}^R)(\text{C}_4\text{H}_4\text{Me}_2\text{-}2,3)(\eta^3\text{-C}_3\text{H}_5)]$ ($\delta_{\text{H}} = 5.65$ ppm, m, 1H, CH; 1.64 ppm, m, 2H, CH_2 -*syn*; -0.83 ppm, d, 2H, $^3J_{\text{HH}} = 8.4$ Hz, CH_2 -*anti*).³⁶ The ^1H NMR spectrum of **4** differs slightly, with six signals integrating to 2:2:2:27:18:18; two resonances are observed for the protons at the 4,5-Cp'' ring positions due to the absence of a mirror plane in **4** compared with parent $[\text{Th}(\text{Cp}^R)_2(\text{Cl})_2]$.²⁴ The $^{13}\text{C}\{^1\text{H}\}$ NMR spectra of $[\text{M}(\text{Cp}^R)_2(\text{Cl})_2]$, **1** and **2** contain the same number of resonances as their respective ^1H NMR spectra, with no resonance observed for the quaternary ring carbon atoms; by contrast resonances for all ring carbon environments are present in the $^{13}\text{C}\{^1\text{H}\}$ NMR spectra of **4** and **5**. The silyl group resonances in the $^{13}\text{C}\{^1\text{H}\}$ NMR spectra of all complexes studied exhibit satellites from coupling to ^{29}Si nuclei, with $^1J_{\text{SiC}} \approx 53$ Hz for Cp'/Cp'' and $^1J_{\text{SiC}} \approx 42$ Hz for hypersilanide silyl groups. The allyl group signals in the $^{13}\text{C}\{^1\text{H}\}$ NMR spectrum of **5** are best compared to those of $[\text{Hf}(\text{Cp}^R)(\text{C}_4\text{H}_4\text{Me}_2\text{-}2,3)(\eta^3\text{-C}_3\text{H}_5)]$ ($\delta_{\text{C}} = 127.87$ ppm, CH; 57.13 ppm, CH_2).³⁶

The $^{29}\text{Si}\{^1\text{H}\}$ NMR spectra of **1**, **2**, **4** and **5** all exhibit signals for the silyl groups of Cp^R rings and the hypersilanide ligands. Most importantly, resonances for the metal-bound silicon atoms were observed for each of these complexes, with the δ_{Si} values of structurally analogous **1** (-80.74 ppm) and **2** (-77.11 ppm) being similar to each other and previously reported Zr and Hf hypersilanide complexes, *e.g.* $[\text{M}(\text{Cp}^*)(\text{Cp})\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ ($\text{M} = \text{Zr}, \delta_{\text{Si}} = -87.30$ ppm;³⁷ $\text{M} = \text{Hf}, \delta_{\text{Si}} = -77.87$ ppm³⁸), $[\text{M}(\text{Cp})_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ ($\text{M} = \text{Zr}, \delta_{\text{Si}} = -85.5$ ppm; $\text{M} = \text{Hf}, \delta_{\text{Si}} = -79.7$ ppm).^{39,40} The substitution of chloride with allyl in **5** leads to a significant upfield shift ($\delta_{\text{Si}} = -108.82$ ppm), highlighting the sensitivity of silanide chemical shifts to the metal coordination environment. Finally, the δ_{Si} value of the metal-bound silicon atom in the Th hypersilanide complex **4** (-66.32 ppm) is significantly downfield to that of the only other known Th(IV) silanide complex, $[\text{Th}(\text{Cp}^R)_3\{\text{Si}(\text{SiMe}_3)_3\}]$ (-108.92 ppm).¹⁹

Single crystal XRD

Complexes **2–6** and $[\text{Hf}(\text{Cp}^R)_2(\text{Cl})_2]$ were studied by single crystal XRD; the solid-state structures of **1**²¹ and other M(IV)



Table 1 ^1H , $^{13}\text{C}\{^1\text{H}\}$ and $^{29}\text{Si}\{^1\text{H}\}$ NMR spectral data for $[\text{M}(\text{Cp}^{\text{R}})_2(\text{Cl})_2]$, **1**, **2**, **4** and **5** at 298 K in C_6D_6

Complex	Environment	$\delta_{\text{H}}/\text{ppm}$	$\delta_{\text{C}}/\text{ppm}$	$\delta_{\text{Si}}/\text{ppm}$
[Zr(Cp') ₂ (Cl) ₂] ²²	<i>SiMe</i> ₃ -Cp'	0.33, s, 18H ^a	0.15 ^a	-6.40
	3,4-Cp'	5.95, m, 4H ^a	115.36 ^a	—
	2,5-Cp'	6.40, m, 4H ^a	126.12 ^a	—
	1-Cp'	—	— ^b	—
[Hf(Cp') ₂ (Cl) ₂] ²²	<i>SiMe</i> ₃ -Cp'	0.32, s, 18H ^a	0.21 ^a	-6.53
	3,4-Cp'	5.89, m, 4H ^a	114.28 ^a	—
	2,5-Cp'	6.31, m, 4H ^a	124.94 ^a	—
	1-Cp'	—	— ^b	—
[Hf(Cp'') ₂ (Cl) ₂] ²⁰	<i>SiMe</i> ₃ -Cp''	0.40, s, 36H	0.29	-6.97
	4,5-Cp''	6.44, m, 4H	119.48	—
	2-Cp''	7.20, m, 2H	142.34	—
	1,3-Cp''	—	— ^b	—
[Th(Cp'') ₂ (Cl) ₂] ²⁴	<i>SiMe</i> ₃ -Cp''	0.37, s, 36H	0.23	-8.83
	4,5-Cp''	6.95, m, 4H	139.86	—
	2-Cp''	7.24, m, 2H	140.47	—
	1,3-Cp''	—	— ^b	—
1 ²¹	<i>SiMe</i> ₃ -Cp'	0.26, s, 18H	0.19	-6.13 ^c
	3,4-Cp'	5.02, m, 2H; 5.82, m, 2H	111.22, 112.46	—
	2,5-Cp'	6.76, m, 2H; 7.67, m, 2H	114.85, 120.46	—
	1-Cp'	—	129.47	—
	Si(<i>SiMe</i> ₃) ₃	0.48, s, 27H	5.22	-6.61 ^c
	Si(<i>SiMe</i> ₃) ₃	—	—	-80.74 ^c
2	<i>SiMe</i> ₃ -Cp'	0.26, s, 18H	0.23	-6.12
	3,4-Cp'	5.00, m, 2H; 5.87, m, 2H	109.99, 112.59	—
	2,5-Cp'	6.59, m, 2H; 7.50, m, 2H	115.00, 119.07	—
	1-Cp'	—	— ^b	—
	Si(<i>SiMe</i> ₃) ₃	0.49, s, 27H	5.43	-5.66
	Si(<i>SiMe</i> ₃) ₃	—	—	-77.11
4	<i>SiMe</i> ₃ -Cp''	0.31, s, 18H; 0.34, s, 18H	0.96, 1.18	-7.49, -7.67
	4,5-Cp''	7.06, m, 2H; 7.46, m, 2H	129.70, 131.12	—
	2-Cp''	7.32, m, 2H	132.76	—
	1,3-Cp''	—	140.17, 141.58	—
	Si(<i>SiMe</i> ₃) ₃	0.60, s, 27H	6.71	-0.71
	Si(<i>SiMe</i> ₃) ₃	—	—	-66.31
5	<i>SiMe</i> ₃ -Cp'	0.16, s, 18H	0.24	-6.20
	3,4-Cp'	4.88, m, 2H; 5.36, m, 2H	102.69, 104.55	—
	2,5-Cp'	5.62, m, 2H; 6.06, m, 2H	105.10, 110.27	—
	1-Cp'	—	115.23	—
	Si(<i>SiMe</i> ₃) ₃	0.52, s, 27H	6.47	-5.29
	Si(<i>SiMe</i> ₃) ₃	—	—	-108.82
	CH(CH ₂) ₂	2.58, d, 4H, ³ J _{HH} = 12.0 Hz	55.86	—
	CH(CH ₂) ₂	4.03, pent, 1H, ³ J _{HH} = 12.0 Hz	112.67	—

^a ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR data originally reported in CDCl_3 but recollected here in C_6D_6 . ^b Not observed/observed by solvent resonance. ^c Not reported previously, collected herein.

precursors $[\text{M}(\text{Cp}^{\text{R}})_2(\text{Cl})_2]$ ^{20,22–24} have all been characterised by this method previously (see Fig. 1 and 2 for depictions of **2–6**, and Table 2 for selected bond distances and angles; see ESI Fig. S43† for a depiction of $[\text{Hf}(\text{Cp}')_2(\text{Cl})_2]$, together with Tables S1–S3† for additional crystallographic data). All complexes exhibit bent metallocene geometries with typical $\text{M}\cdots\text{Cp}^{\text{R}}_{\text{cent}}$ distances for the metals and Cp^{R} ligands present;¹⁸ $[\text{Hf}(\text{Cp}')_2(\text{Cl})_2]$ exhibits a similar solid-state structure to both its Zr congener²² and the related Hf complex $[\text{Hf}(\text{Cp}'')_2(\text{Cl})_2]$,²⁰ and as its metrical parameters are unremarkable we do not comment on this structure further.

The solid-state structure of **2** is analogous to that of **1**, though the bond distances are all shorter in the former complex, as exemplified by the Hf–Si distance (2.7901(9) Å) being shorter than the corresponding Zr–Si bond reported for **1** (2.833(3) Å);²¹ this is in accord with the greater covalent radii of Zr (1.54 Å) vs. Hf (1.52 Å).⁴¹ The structure of **3** is comparable

to that of **2**, though the additional steric buttressing provided by the additional silyl group in the Cp'' ring leads to a slightly longer Hf–Si bond (2.8057(8) Å) and a more obtuse $\text{Cp}^{\text{R}}_{\text{cent}}\cdots\text{Hf}\cdots\text{Cp}^{\text{R}}_{\text{cent}}$ angle (131.70(5)° *c.f.* 128.37(5)° for **2**). Similarly, the Hf–Si bond (2.871(2) Å) in **5** is elongated due to the additional steric bulk of the allyl ligand compared to the chloride in **2**; this distance is still shorter than that previously seen for $[\text{Hf}(\text{Cp}^*)(\text{Cp})\{\text{Si}(\text{SiMe}_3)_3(\text{Cl})\}]$ (2.881(4) Å), due to increased steric buttressing by Cp* vs. Cp''.³⁸ The Hf $\cdots\text{C}_{\text{allyl}}$ distances of **5** (range: 2.426(5)–2.555(5) Å) are relatively long in comparison with previously reported Hf allyl complexes,¹⁸ *e.g.* $[\text{Hf}(\text{Cp}'')(\text{C}_4\text{H}_4\text{Me}_2-2,3)(\eta^3\text{-C}_3\text{H}_5)]$ (Hf $\cdots\text{C}_{\text{allyl}}$ range: 2.409(6)–2.452(6) Å);³⁶ this is in accord with the relatively weak binding of the η^3 -allyl in **5** indicated by solution ^1H NMR studies (see above). The Th–Si distance of **4** (3.1053(13) Å) is shorter than that of $[\text{Th}(\text{Cp}')_3\{\text{Si}(\text{SiMe}_3)_3\}]$ (3.1191(8) Å),¹⁹ as expected from the decreased combined steric requirements of one halide and



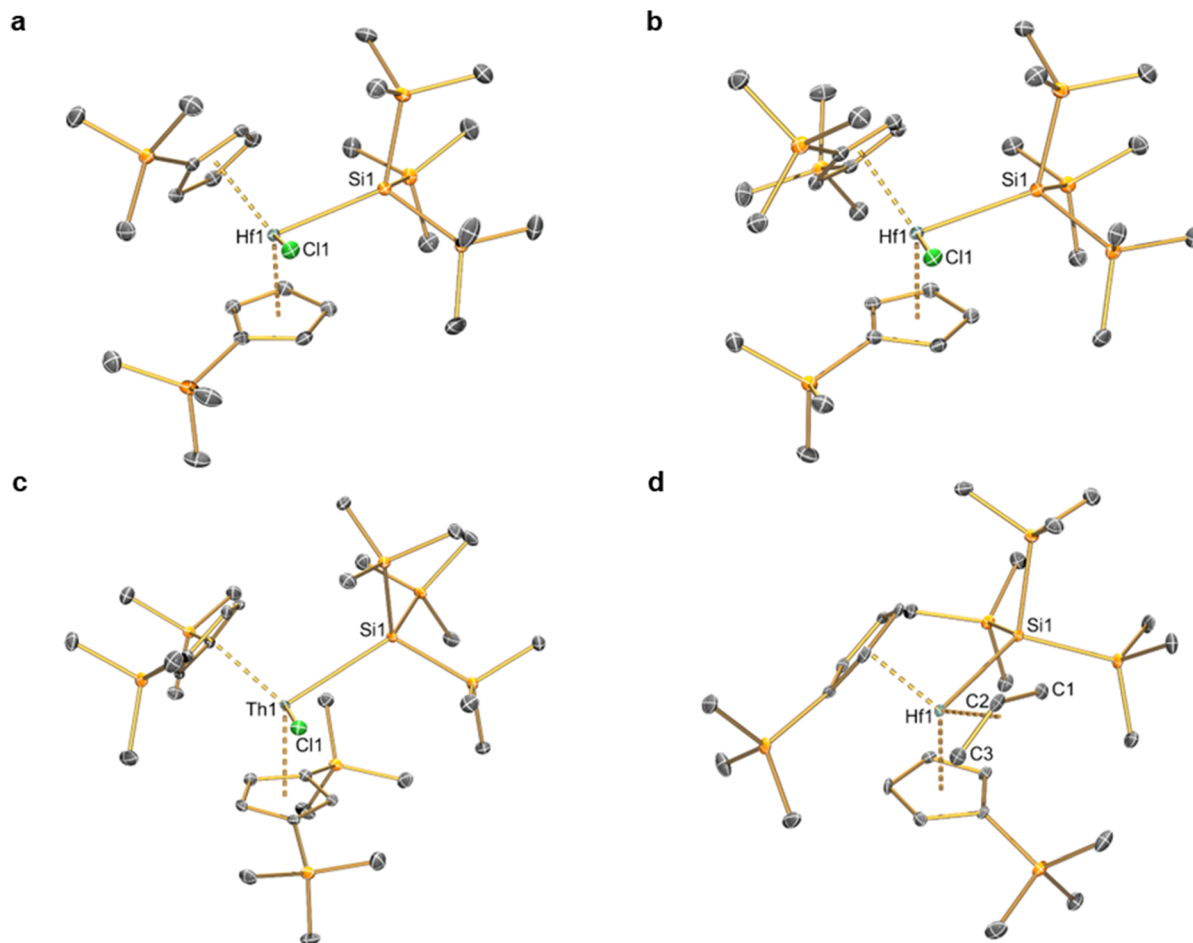


Fig. 1 Single crystal XRD structure of (a) $[\text{Hf}(\text{Cp}')_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ (**2**), (b) $[\text{Hf}(\text{Cp}'')(\text{Cp}')\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]\cdot 0.5\text{C}_5\text{H}_{12}$ (**3·0.5C₅H₁₂**), (c) $[\text{Th}(\text{Cp}'')_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ (**4**) and (d) $[\text{Hf}(\text{Cp}')_2\{\text{Si}(\text{SiMe}_3)_3\}(\eta^3\text{-C}_3\text{H}_5)]$ (**5**) with selective atom labelling. Displacement ellipsoids set at 30% probability level and hydrogen atoms, lattice solvent and modelled disorder components omitted for clarity.

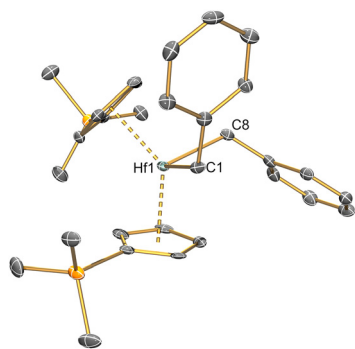


Fig. 2 Single crystal XRD structure of $[\text{Hf}(\text{Cp}')_2(\text{CH}_2\text{Ph})_2]$ (**6**) with selective atom labelling. Displacement ellipsoids set at 30% probability level and hydrogen atoms omitted for clarity.

two Cp'' ligands in the former complex vs. the tris- Cp' motif in the latter. In 1–5 all M–Si distances are longer than the sum of single bond covalent radii for the respective metal and Si (M = Zr, 2.70 Å; Hf, 2.68 Å; Th, 2.91 Å).⁴¹ The presence of a bulky hypersilanide ligand in 1–5 does not lead to significant

changes in M–Cl and $\text{M}\cdots\text{Cp}^{\text{R}}_{\text{cent}}$ distances compared to parent $[\text{M}(\text{Cp}^{\text{R}})_2(\text{Cl})_2]$.^{20,22–24} Finally, the mean Hf–C distances (2.275(8) Å) and Hf–C– C_{ipso} angles (121.4(4) and 124.9(4)°) of **6** are comparable to previously reported Hf benzyl complexes,¹⁸ with the most closely related example being $[\text{Hf}(\text{Cp})_2(\text{CH}_2\text{Ph})_2]$ (mean Hf–C = 2.293(10) Å, Hf–C– C_{ipso} angles = 120.1(5) and 123.3(5)°).⁴²

Quantum chemical calculations

We performed restricted DFT calculations on 1–6, in order to probe the M–Si and M–C σ -bonds in these complexes further (Table 3). As geometry optimisations effectively reproduced the metrical parameters observed by single crystal XRD we posit that these models are representative of the electronic structures of 1–6; although there are minor variations in the steric bulk provided by ancillary ligands, qualitative comparisons can be made (see ESI Tables S4–S9† for atomic coordinates of geometry-optimised structures). For 1–5 the calculated M–Si distances are longer than those determined experimentally but follow the same trend, with the M–Si bond of **2** shorter than that of **1** and the M–Si bond of **5** being longer than that of **1**



Table 2 Selected bond lengths (Å), mean M...Cp^R_{cent} distances (Å) and Cp^R_{cent}...M...Cp^R_{cent} and Si–M–X angles (°) for 1–6

Parameter	1 ²¹	2	3-0.5C ₅ H ₁₂	4	5	6 ^a
M–Si	2.833(3)	2.7901(9)	2.8057(8)	3.1053(13)	2.871(2)	—
M–Cl	2.430(2)	2.4041(7)	2.3977(7)	2.6118(11)	—	—
Hf–C ^b	—	—	—	—	2.426(5), 2.501(5), 2.555(5)	2.281(6), 2.268(6)
Mean M...Cp ^R _{cent}	2.21(1)	2.190(3)	2.198(2)	2.520(4)	2.211(4)	2.198(4)
Si–M–Cl	91.5(1)	91.20(3)	91.10(2)	89.25(6)	—	—
Cp ^R _{cent} ...M...Cp ^R _{cent}	128.3(4)	128.37(5)	131.70(5)	124.33(12)	127.78(10)	131.68(11)

^a Two independent molecules in asymmetric unit, parameters shown for one of these molecules. ^b Three allyl Hf...C distances for 5 and two Hf–C alkyl bond lengths for 6.

Table 3 Selected computed DFT, NBO and QTAIM data for the M–Si and M–C σ -bonds in 1–6

Entry ^a	Bond	Bond length and index ^{b,c}		Charges			Natural bond orbital (NBO) analyses				QTAIM ^g			
		Bond Length	Bond Index	q_M^d	q_{Si}^e	q_C^f	M [%]	Si/C [%]	M s/p/d/f [%]	C/Si ns/np [%]	$\rho(r)$	$\nabla^2\rho(r)$	$H(r)$	$\epsilon(r)$
1	Zr–Si	2.8787	0.89	0.66	–0.28	—	32	68	17/0/83/0	36/64	0.05	–0.03	–0.02	0.05
2	Hf–Si	2.8483	0.89	0.69	–0.17	—	29	71	27/0/73/0	37/63	0.05	–0.04	–0.02	0.04
3	Hf–Si	2.8652	0.88	0.71	–0.16	—	28	72	29/0/71/0	37/63	0.05	–0.03	–0.02	0.05
4	Th–Si	3.1759	0.84	1.87	–0.68	—	19	81	34/0/54/12	39/61	0.07	0.07	–0.01	0.01
5	Hf–Si	2.9676	0.75	0.77	–0.17	—	29	71	27/0/73/0	37/63	0.04	0.01	–0.01	0.09
6	Hf–C	2.3232	0.65	1.15	—	–0.99	18	82	27/0/73/0	23/77	0.08	0.08	–0.04	0.04

^a All molecules geometry-optimised without symmetry constraints at the LDA VWN BP86 TZP/ZORA level. ^b Calculated bond of interest. ^c Nalewajski–Mrozek bond indices. ^d Multipole Derived Charges (MDC-q) on M. ^e MDC-q charge on Si. ^f MDC-q charge on C. ^g Quantum Theory of Atoms in Molecules (QTAIM) topological electron density [$\rho(r)$], Laplacian [$\nabla^2\rho(r)$], electronic energy density [$H(r)$], and ellipticity [$\epsilon(r)$] 3, –1 bond critical point data.

and 2, but shorter than the Th–Si bond of 4. The Nalewajski–Mrozek bond indices of 1–3 (range: 0.88–0.89) are essentially identical, with the Th–Si bond of 4 (0.84) lower than these values but higher than the Hf–Si bond of 5 (0.75). The Multipole-Derived Charges (MDC-q) vary markedly for different metals. The q_{Si} values for 2, 3 and 5 are consistently –0.16 to –0.17, with higher values for 1 (–0.28) and 4 (–0.68), and whilst the q_M values are similar for the d-block complexes 1–3 and 5 (range: 0.66–0.77), the q_{Th} value for 4 is higher (1.78), showing a higher polarity of M–Si bonds for M = Th vs. Zr and Hf. The Hf–Si bond polarity is higher for 5 than for 2 and 3, showing that the allyl ligand is a stronger donor than chloride. The q_{Hf} of 6 is 1.15 and the q_C is –0.99, indicating that there is poorer orbital energy matching between Hf and C than between Hf and Si. The Hf–C bonds in 6 are therefore more polar than the Hf–Si bonds in 2, 3 and 5, but are more covalent than the corresponding Th–Si bond in 4.

Natural Bond Orbital (NBO) analyses of the M–Si and M–C bonds in 1–6 show that these are the Highest Occupied Molecular Orbitals (HOMOs) in all cases (see Fig. 3 for Kohn–Sham Molecular Orbital depictions for 1–6, and ESI Fig. S44–S49† for other selected Kohn–Sham Molecular Orbitals and NBOs; N.B. for 6 the HOMO–1 is also a Hf–C σ -bond). The Zr–Si bond in 1 has the most metal character (32%), followed by the Hf–Si bonds in 2, 3 and 5 (28–29%), with the Th–Si bond in 4 showing the least metal contribution (19%) as expected.¹⁶ The Hf–C bonds of 6 have a lower Hf contribution (18%) than any of the Hf–Si bonds, but these are all remarkably invariant

to each other with respect to orbital character (27–29% 6s : 71–73% 5d); the Zr–Si bond in 1 shows a greater metal nd -orbital component (17% 5s : 83% 4d), whilst the Th–Si bond in 4 has the most ns , the least nd , and some 5f contribution (34% 7s : 54% 6d : 12% 5f). The Si orbital contributions to the M–Si bonds in 1–5 do not vary markedly (36–39% 3s : 61–64% 3p), whilst there is a greater np character from C in the Hf–C bonds of 6 (2s 23% : 2p 77%).

Quantum Theory of Atoms in Molecules (QTAIM) analyses of 1–6 reveals highly polarised (range of topological electron densities, $\rho(r)$: 0.05–0.08) and cylindrical (range of bond ellipticities, $\epsilon(r)$: 0.01–0.09) M–Si and M–C σ -bonds. Electronic energy densities, $H(r)$, are all negative as expected and for heavy atoms the Laplacian values ($\nabla^2\rho(r)$, range: 0.04–0.07) are insignificant, as we have commented on previously in DFT analyses of diamagnetic f-block silanide complexes.^{19,43,44} The QTAIM parameters extracted are consistent with the NBO data and bond indices and a visual evaluation of the HOMOs, which all indicate that the Th–Si bond of 4 is less covalent than the Zr–Si bond of 1 and the Hf–Si bonds of 2, 3 and 5. Conversely, the metal-bound silanide resonance in the ²⁹Si NMR spectrum of 4 (–66.31 ppm) is downfield to that of 1 (–80.74 ppm) and 2 (–77.11 ppm), which would indicate lower polarity of the Th–Si bond. This discrepancy could be attributed to Heavy Atom on Light Atom (HALA) spin–orbit effects;⁴⁵ this is also in-line with the larger $\rho(r)$ value of 4 compared to 1–3 perhaps reflecting the more diffuse nature of 6d orbitals, but we note the smaller $H(r)$ value for 4 vs. 1–3. A



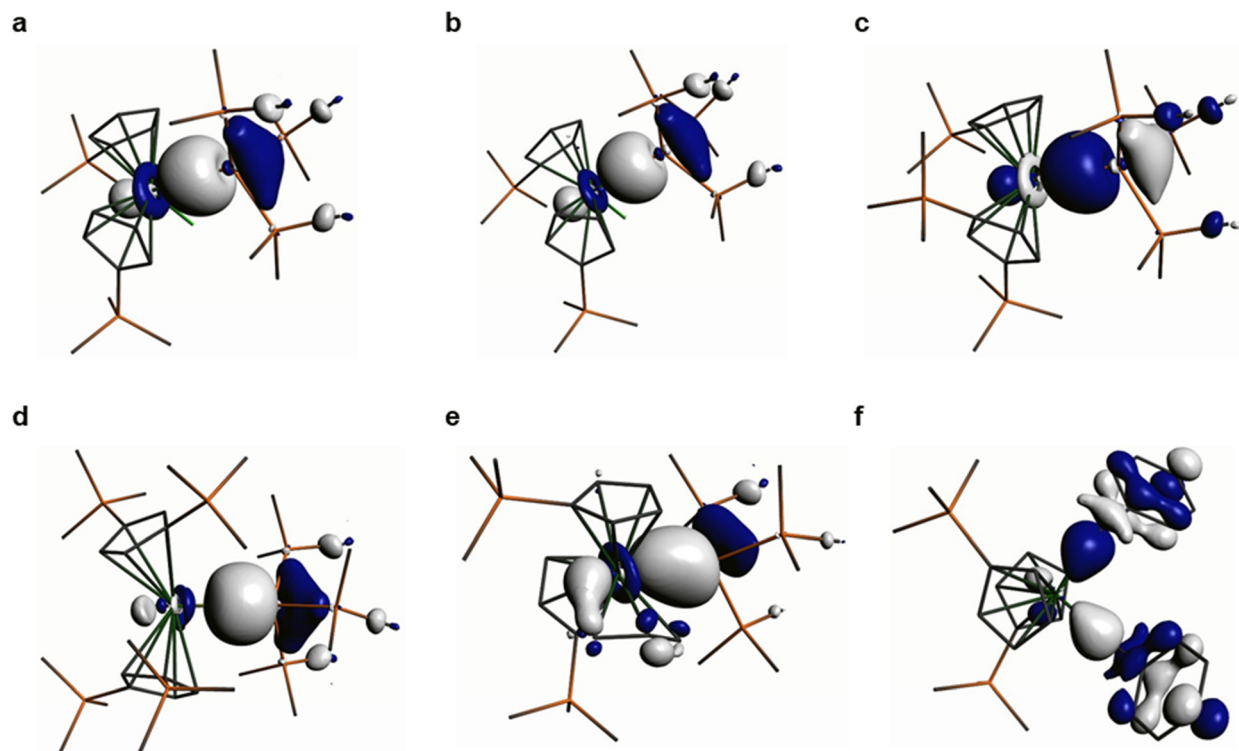


Fig. 3 Kohn-Sham Highest Occupied Molecular Orbitals (HOMOs) representing the M-Si or M-C σ -bonds in 1–6. (a) 1 (–4.866 eV); (b) 2 (–4.919 eV); (c) 3 (–4.884 eV); (d) 4 (–4.616 eV); (e) 5 (–4.571 eV); (f) 6 (–4.800 eV). Hydrogen atoms are omitted for clarity.

quantitative comparison of the DFT-calculated parameters for **4** with the previously reported Th(IV) hypersilanide complex [Th(Cp^R)₃{Si(SiMe₃)₃}] is not possible as these were performed at a different level of theory,¹⁹ but qualitatively the Th-Si bond of **4** appears to be more polarised and have a lower metal 6d orbital contribution. This highlights that the replacement of a Cp^R ligand with a chloride ligand in the Th coordination sphere can have a significant effect on the Th-Si bond, though we caution against making firm conclusions as the calculated Th-Si bond for **4** (3.1759 Å) is 2.3% longer than the distance determined by single crystal XRD (3.1053(13) Å), whereas for [Th(Cp^R)₃{Si(SiMe₃)₃}] there was only a 1.6% discrepancy between theory and experiment (3.071 Å calculated vs. 3.1191(8) Å measured).¹⁹

Conclusion

We have used salt metathesis protocols to synthesise a series of M(IV) hypersilanide complexes for M = Hf and Th, supported by two substituted cyclopentadienyls and one chloride or allyl ancillary ligand, in order to compare their M-Si bonds with each other and previously reported similar Th(IV)¹⁹ and Zr(IV)²¹ hypersilanide complexes. Attempts to prepare neutral Th(III) or cationic M(IV) (M = Hf, Th) hypersilanide analogues of these complexes by respective reduction or anion abstraction protocols were unsuccessful under the reaction conditions investigated. However, this work has provided the second structurally

authenticated example of a Th-Si bond, and has allowed an extended solid-state structural and ²⁹Si{¹H} NMR spectroscopic comparison of group 4 metal and Th(IV) silanide complexes. We have also shown the importance of the steric bulk of supporting silyl-substituted cyclopentadienyl ligands in dictating the reaction outcome when installing a hypersilanide ligand, due to differences in M(IV) ion size, Lewis acidity and covalency between these metals. DFT calculations of structurally similar M(IV) hypersilanide complexes for M = Zr, Hf and Th confirmed as expected that the highly polarised Th-Si bonds are less covalent than Zr-Si and Hf-Si bonds, which exhibit similar covalency to each other.

Experimental

General methods

All manipulations were performed under strict anaerobic conditions using argon as the inert gas, by using a combination of Schlenk line and glove box techniques. HPLC grade pentane, toluene and THF were passed through alumina columns in a solvent purification system. Benzene, 1,4-dioxane and DME were dried by refluxing over potassium for 24 h followed by distillation. All solvents were degassed under dynamic vacuum and were stored over potassium mirrors (benzene pentane, toluene) or 4 Å molecular sieves (1,4-dioxane, DME, THF) before use. C₆D₆ was dried by refluxing over potassium for 2 days in a J. Young tap-appended ampoule before trap-to-trap



distillation and three freeze–pump–thaw cycles to degas before use. $[\text{Zr}(\text{Cp}')_2(\text{Cl})_2]$ ($\text{M} = \text{Zr}, \text{Hf}$),²² $[\text{M}(\text{Cp}'')_2(\text{Cl})_2]$ ($\text{M} = \text{Hf}, \text{Th}^{24}$), $\text{K}\{\text{Si}(\text{SiMe}_3)_3\}$,²⁵ KCH_2Ph ,²⁸ $[\text{CPh}_3][\text{Al}\{\text{OC}(\text{CF}_3)_3\}_4]$,²⁹ $[\text{NEt}_3\text{H}][\text{BPh}_4]$ ³¹ and KC_8 ³² were prepared by previously reported procedures, whilst complex **1** was prepared by adapting the literature method.²¹ Details of the revised synthesis and data obtained for **1** are provided below, together with synthetic procedures and experimental data obtained for **2–6**. All other starting materials were purchased from commercial sources and were used as received. **Caution:** *Thorium from natural sources is a weak α -emitter; it is recommended that compounds of this element are only manipulated by trained personnel in designated radiochemical laboratories, and that α -counting equipment is routinely available for monitoring the radiochemical hazard.*

NMR spectra were recorded on samples in J. Young tap-appended NMR tubes using a Bruker Avance III 400 spectrometer operating at 400.07 (^1H), 100.60 (^{13}C) or 79.48 (^{29}Si) MHz. ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR spectra were internally referenced to the residual solvent resonance of C_6D_6 , whilst $^{29}\text{Si}\{^1\text{H}\}$ NMR spectra were externally referenced to tetramethylsilane; chemical shifts are provided in ppm. NMR data was analysed using the MestReNova software suite.⁴⁶ ATR-IR spectra of microcrystalline powders were collected on a Bruker Alpha 2 spectrometer. Elemental analyses were carried out by Mr Martin Jennings and Mrs Anne Davies using a Flash 2000 elemental analyser at the Microanalytical service, Department of Chemistry, the University of Manchester. Results generally showed reasonable agreement with expected compositions, though low carbon compositions were reproducibly measured for **4**. This is a common occurrence for early metal complexes with high silicon content, including for f-block Cp'' complexes, where it has been attributed to the formation of silicon carbide, leading to incomplete combustion;⁴⁷ the observation of low carbon values in elemental analyses of $[\text{Th}(\text{Cp}')_3\{\text{Si}(\text{SiMe}_3)_3\}]$ was previously ascribed to the same phenomenon.¹⁹

Crystallographic methods

Crystals of **1–6** were dispersed in fomblin and single crystals were selected and transferred to MicroMounts™ on a goniometer head under a cryostream for exposure to X-rays. Crystallographic data for $[\text{Hf}(\text{Cp}')_2(\text{Cl})_2]$, **2**, **5** and **6** were obtained using an Oxford Diffraction Xcalibur diffractometer equipped with an Agilent Atlas CCD detector with graphite-monochromated $\text{Mo K}\alpha$ ($\lambda = 0.71073 \text{ \AA}$) radiation, whilst data for **3-0.5C₅H₁₂** and **4** were collected using a Rigaku FR-X diffractometer equipped with a HyPix 6000HE photon counting pixel array detector with mirror-monochromated $\text{Cu K}\alpha$ (1.54184 \AA) radiation. Intensities were integrated from data recorded on 0.5° (**5**) or 1° ($[\text{Hf}(\text{Cp}')_2(\text{Cl})_2]$, **2**, **3-0.5C₅H₁₂**, **4** and **6**) frames by ω rotation. Cell parameters were refined from the observed positions of all strong reflections in each data set. A Gaussian grid face-indexed with a beam profile was applied for all structures.⁴⁸ The SHELXT program was used to solve all structures.⁴⁹ Datasets were refined by full-matrix least-squares on all unique F^2 values,⁴⁹ with anisotropic displacement para-

eters for all non-hydrogen atoms, and with constrained riding hydrogen geometries; $U_{\text{iso}}(\text{H})$ was set at 1.2 (1.5 for methyl groups) times U_{eq} of the parent atom. The largest features in final difference syntheses were close to heavy atoms and were of no chemical significance. The CrysAlisPro program⁴⁸ was used for control and integration, and the SHELX suite^{49,50} was employed through the OLEX2 program⁵¹ for structure solution and refinement. The programs ORTEP-3⁵² and POV-Ray⁵³ were used in combination to generate depictions of **1–6**.

Computational methods

The Amsterdam Density Functional (ADF) program (2017 version) was used to perform restricted DFT calculations using standard convergence criteria.^{54,55} Atomic coordinates determined from single crystal XRD were used as the start points for geometry optimisations, which were performed with no constraints applied. Slater-type orbital TZP polarisation all-electron basis sets from the Dirac and zeroth order regular approximation (ZORA) and triple-zeta with polarisation (TZP) database of the Amsterdam Density Functional (ADF) package were used for DFT geometry optimisations. Scalar relativistic approaches (spin-orbit neglected) were used within the ZORA Hamiltonian to include relativistic effects.^{56–58} The local density approximation (LDA) with a correlation potential was used in all calculations.⁵⁹ Generalised gradient approximation corrections were performed using the functionals of Becke and Perdew.^{60,61} NBO analysis was carried out using the NBO v6.0 program.⁶² QTAIM analysis^{63,64} was performed within the ADF suite. The Kohn–Sham MOs and NBOs were visualised using ADFView.

Synthetic procedures

General procedure for the synthesis of $[\text{M}(\text{Cp}^{\text{R}})_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ (1–4**).** A Schlenk flask was charged with $[\text{M}(\text{Cp}^{\text{R}})_2(\text{Cl})_2]$ ($\text{M} = \text{Zr}, \text{Hf}, \text{Cp}^{\text{R}} = \text{Cp}' = \{\text{C}_5\text{H}_4\text{SiMe}_3\}; \text{M} = \text{Hf}, \text{Th}, \text{Cp}^{\text{R}} = \text{Cp}'' = \{\text{C}_5\text{H}_3(1,3\text{-SiMe}_3)\}$), cooled to -78°C and dissolved in toluene (15 mL mmol^{-1}). $\text{K}\{\text{Si}(\text{SiMe}_3)_3\}$ (1 eq.) was dissolved in toluene (10 mL mmol^{-1}) and added dropwise to the cooled $[\text{M}(\text{Cp}^{\text{R}})_2(\text{Cl})_2]$. The red ($\text{M} = \text{Zr}$) or orange ($\text{M} = \text{Hf}, \text{Th}$) suspension was warmed to room temperature and stirred for 1 hour. All volatiles were subsequently removed *in vacuo*, and the solids were extracted with pentane (40 mL mmol^{-1}). Concentration and storage of solutions at -25°C led to the formation of needles of $[\text{M}(\text{Cp}^{\text{R}})_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$.

Preparation of $[\text{Zr}(\text{Cp}')_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ (1**).** Modified from previous literature method,²¹ prepared according to the general procedure above with $[\text{Zr}(\text{Cp}')_2(\text{Cl})_2]$ (0.873 g, 2 mmol) and $\text{K}\{\text{Si}(\text{SiMe}_3)_3\}$ (0.574 g, 2 mmol); $[\text{Zr}(\text{Cp}')_2\{\text{Si}(\text{SiMe}_3)_3\}(\text{Cl})]$ (**1**) was obtained as red needles (0.792 g, 1.22 mmol, 61%). Anal. Calcd for $\text{C}_{25}\text{H}_{53}\text{ClSi}_6\text{Zr}$: C, 46.28; H, 8.23. Found: C, 46.69; H, 8.56. ^1H NMR (400.07 MHz, C_6D_6 , 298 K): $\delta = 0.26$ (s, 18H, $\text{Cp}'\text{-Si}(\text{CH}_3)_3$), 0.48 (s, 27H, $\text{Si}\{\text{Si}(\text{CH}_3)_3\}_3$), 5.02 (m, 2H, 3,4- $\text{Cp}'\text{-H}$), 5.82 (m, 2H, 3,4- $\text{Cp}'\text{-H}$), 6.76 (m, 2H, 2,5- $\text{Cp}'\text{-H}$), 7.67 (m, 2H, 2,5- $\text{Cp}'\text{-H}$). $^{13}\text{C}\{^1\text{H}\}$ NMR (100.60 MHz, C_6D_6 , 298 K): $\delta = 0.19$ ($J_{\text{SiC}} = 53.8 \text{ Hz}$, $\text{Cp}'\text{-Si}(\text{CH}_3)_3$), 5.22 ($J_{\text{SiC}} = 42.1$



Hz, Si{Si(CH₃)₃}₃, 111.22 (3,4-Cp'-CH), 112.46 (3,4-Cp'-CH), 114.85 (2,5-Cp'-CH), 120.46 (2,5-Cp'-CH), 129.47 (1-Cp'-C). ²⁹Si{¹H} NMR (79.48 MHz, C₆D₆, 298 K): δ = -80.74 (SiZr), -6.61 (Si{Si(CH₃)₃}₃), -6.13 (Cp'-Si(CH₃)₃). ATR-IR $\tilde{\nu}/\text{cm}^{-1}$: 2949 (m), 2890 (m), 1445 (w), 1406 (m), 1393 (m), 1375 (m), 1245 (s), 1172 (m), 1046 (m), 906 (m), 824 (s), 808 (s), 754 (s), 678 (s), 623 (s), 427 (s).

Preparation of [Hf(Cp')₂{Si(SiMe₃)₃}Cl] (2). Prepared according to the general procedure with [Hf(Cp')₂(Cl)₂] (0.524 g, 1 mmol) and K{Si(SiMe₃)₃} (0.287 g, 1 mmol); [Hf(Cp')₂{Si(SiMe₃)₃}Cl] was obtained as orange needles (0.519 g, 0.70 mmol, 70%). Anal. Calcd for C₂₅H₅₃ClHfSi₆: C, 40.79; H, 7.26. Found: C, 40.83; H, 7.45. ¹H NMR (400.07 MHz, C₆D₆, 298 K): δ = 0.26 (s, 18H, Cp'-Si(CH₃)₃), 0.49 (s, 27H, Si{Si(CH₃)₃}₃), 5.00 (m, 2H, 3,4-Cp'-H), 5.87 (m, 2H, 3,4-Cp'-H), 6.59 (m, 2H, 2,5-Cp'-H), 7.50 (m, 2H, 2,5-Cp'-H). ¹³C{¹H} NMR (100.60 MHz, C₆D₆, 298 K): δ = 0.23 (¹J_{SiC} = 53.3 Hz, Cp'-Si(CH₃)₃), 5.43 (¹J_{SiC} = 42.3 Hz, Si{Si(CH₃)₃}₃), 109.99 (3,4-Cp'-CH), 112.59 (3,4-Cp'-CH), 115.00 (2,5-Cp'-CH), 119.07 (2,5-Cp'-CH), 1-Cp'-C not observed. ²⁹Si{¹H} NMR (79.48 MHz, C₆D₆, 298 K): δ = -77.11 (SiHf), -6.12 (Cp'-Si(CH₃)₃), -5.66 (Si{Si(CH₃)₃}₃). ATR-IR $\tilde{\nu}/\text{cm}^{-1}$: 2951 (m), 2892 (m), 1438 (w), 1406 (m), 1393 (m), 1376 (m), 1245 (s), 1178 (m), 1046 (m), 908 (m), 818 (s), 812 (s), 754 (s), 678 (s), 623 (s), 427 (s).

Preparation of [Hf(Cp'')(Cp')Si(SiMe₃)₃]Cl] (3). Prepared according to the general procedure with [Hf(Cp'')(Cl)₂] (0.524 g, 1 mmol) and K{Si(SiMe₃)₃} (0.287 g, 1 mmol); [Hf(Cp'')(Cp')Si(SiMe₃)₃]Cl] was obtained as orange needles in <1% crystalline yield, precluding the collection of additional characterisation data.

Preparation of [Th(Cp'')(Cp')Si(SiMe₃)₃]Cl] (4). Prepared according to the general procedure with [Th(Cp'')(Cl)₂] (1.805 g, 2.5 mmol) and K{Si(SiMe₃)₃} (0.717 g, 2.5 mmol); 4 was obtained as orange needles (1.756 g, 1.88 mmol, 75%). Anal. Calcd for C₃₁H₆₉ClSi₈Th: C, 39.86; H, 7.45. Found: C, 38.89; H, 7.51. ¹H NMR (400.07 MHz, C₆D₆, 298 K): δ = 0.31 (s, 18H, Cp''-Si(CH₃)₃), 0.34 (s, 18H, Cp'-Si(CH₃)₃), 0.60 (s, 27H, Si{Si(CH₃)₃}₃), 7.06 (m, 2H, 4,5-Cp''-H), 7.32 (m, 2H, 2-Cp''-H), 7.46 (m, 2H, 4,5-Cp''-H). ¹³C{¹H} NMR (100.60 MHz, C₆D₆, 298 K): δ = 0.96 (¹J_{SiC} = 52.7 Hz, Cp''-Si(CH₃)₃), 1.18 (¹J_{SiC} = 52.1 Hz, Cp'-Si(CH₃)₃), 6.71 (¹J_{SiC} = 42.6 Hz, Si{Si(CH₃)₃}₃), 129.70 (4,5-Cp''-CH), 131.12 (4,5-Cp''-CH), 132.76 (2-Cp''-CH), 140.17 (1,3-Cp''-C), 141.58 (1,3-Cp''-C). ²⁹Si{¹H} NMR (79.48 MHz, C₆D₆, 298 K): δ = -66.32 (SiTh), -7.67 (Cp''-Si(CH₃)₃), -7.50 (Cp'-Si(CH₃)₃), -0.71 (Si{Si(CH₃)₃}₃). ATR-IR $\tilde{\nu}/\text{cm}^{-1}$: 2951 (m, C-H stretch), 2894 (w, C-H stretch), 1436 (w), 1407 (w), 1243 (s), 1077 (s), 1021 (m), 917 (m), 820 (s), 797 (s), 750 (s), 691 (s), 678 (s), 637 (s), 619 (s), 469 (s), 413 (m).

Preparation of [Hf(Cp')₂{Si(SiMe₃)₃}(η³-C₃H₅)] (5). A solution of 2.0 M Mg(C₃H₅)Cl in THF (0.6 mL, 1.2 mmol) was added dropwise *via* a glass syringe with a stainless steel Luer lock needle to a Schlenk flask containing a solution of 2 (0.736 g, 1 mmol) in toluene (20 mL). The yellow reaction mixture was allowed to stir overnight. Volatiles were removed *in vacuo* and pentane (30 mL) and 1,4-dioxane (3 mL) was added and the reaction mixture was stirred for 1 h in order to form MgCl₂·1,4-

dioxane. Volatiles were removed *in vacuo* to give a bright yellow powder, which was extracted with pentane (40 mL). Filtration, concentration and storage of the resultant bright yellow solution at -25 °C led to the formation of yellow blocks of 5 (0.466 g, 0.63 mmol, 63%). Anal. Calcd for C₂₈H₅₈ClHfSi₆: C, 45.34; H, 7.88. Found: C, 45.04; H, 8.16. ¹H NMR (400.07 MHz, C₆D₆, 298 K): δ = 0.16 (s, 18H, Cp'-Si(CH₃)₃), 0.52 (s, 27H, Si{Si(CH₃)₃}₃), 2.58 (d, 4H, ³J_{HH} = 12.0 Hz, CH(CH₂)₂), 4.03 (pent, 1H, ³J_{HH} = 12.0 Hz, CH(CH₂)₂), 4.88 (m, 2H, 3,4-Cp'-H), 5.36 (m, 2H, 3,4-Cp'-H), 5.62 (m, 2H, 2,5-Cp'-H), 6.06 (m, 2H, 2,5-Cp'-H). ¹³C{¹H} NMR (100.60 MHz, C₆D₆, 298 K): δ = 0.24 (¹J_{SiC} = 53.3 Hz, Cp'-Si(CH₃)₃), 6.47 (¹J_{SiC} = 41.1 Hz, Si{Si(CH₃)₃}₃), 55.86 (CH(CH₂)₂), 102.69 (3,4-Cp'-CH), 104.55 (3,4-Cp'-CH), 105.10 (2,5-Cp'-CH), 110.27 (2,5-Cp'-CH), 112.67 (CH(CH₂)₂), 115.23 (1-Cp'-C). ²⁹Si{¹H} NMR (79.48 MHz, C₆D₆, 298 K): δ = -108.82 (SiHf), -6.20 (Cp'-Si(CH₃)₃), -5.29 (Si{Si(CH₃)₃}₃). ATR-IR $\tilde{\nu}/\text{cm}^{-1}$: 2951 (m, C-H stretch), 2897 (w, C-H stretch), 1531 (w, allyl stretch) 1445 (w), 1405 (w), 1376 (w), 1315 (w), 1241 (s), 1167 (m), 1044 (m), 904 (m), 824 (s), 799 (s), 752 (s), 667 (s), 620 (s), 423 (m).

Preparation of [Hf(Cp')₂(CH₂Ph)₂] (6). Toluene (10 mL) was added *via* a stainless steel cannula to a Schlenk flask charged with 2 (0.368 g, 0.5 mmol) and KCH₂Ph (0.065 g, 0.5 mmol) in toluene (20 mL). The orange reaction mixture was allowed to stir for 2 h. Volatiles were removed *in vacuo* to give an orange powder, which was extracted with pentane (25 mL). Filtration, concentration and storage of the resultant orange solution at -25 °C led to the formation of yellow blocks of 6 together with other reaction products, precluding the collection of additional characterisation data.

Conflicts of interest

The authors declare no conflicts of interest.

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References

- 1 *Organosilicon Compounds: Theory and Experiment (Synthesis)*, Ed. V. Y. Lee, Academic Press, Elsevier Inc., Amsterdam, 2017.
- 2 R. Waterman, P. G. Hayes and T. D. Tilley, *Acc. Chem. Res.*, 2007, **40**, 712–719.



- 3 W. Gao, X. Zhang, X. Xie and S. Ding, *Chem. Commun.*, 2020, **56**, 2012–2015.
- 4 A. Walczak, H. Stachowiak, G. Kurpiak, J. Kaźmierczak, G. Hreczycho and A. R. Stefankiewicz, *J. Catal.*, 2019, **373**, 139–146.
- 5 V. Srinivas, Y. Nakajima, K. Sato and S. Shimada, *Org. Lett.*, 2018, **20**, 12–15.
- 6 X. Chen and C. Liang, *Catal. Sci. Technol.*, 2019, **9**, 4785–4820.
- 7 B. L. L. Réant, S. T. Liddle and D. P. Mills, *Chem. Sci.*, 2020, **11**, 10871–10886.
- 8 I. Castillo and T. D. Tilley, *Organometallics*, 2001, **20**, 5598–5605.
- 9 A. D. Sadow and T. D. Tilley, *J. Am. Chem. Soc.*, 2005, **127**, 643–656.
- 10 F. Pan, J. Zhang, H. L. Chen, Y. H. Su, C. L. Kuo, Y. H. Su, S. H. Chen, K. J. Lin, P. H. Hsieh and W. S. Hwang, *Materials*, 2016, **9**, 417.
- 11 J. T. White, A. T. Nelson, J. T. Dunwoody, D. D. Byler, D. J. Safarik and K. J. McClellan, *J. Nucl. Mater.*, 2015, **464**, 275–280.
- 12 T. L. Wilson, E. E. Moore, D. Adorno Lopes, V. Kocovski, E. Sooby Wood, J. T. White, A. T. Nelson, J. W. McMurray, S. C. Middleburg, P. Xu and T. M. Besmann, *Adv. Appl. Ceram.*, 2018, **117**, S76–S81.
- 13 U. E. Humphrey and M. U. Khandaker, *Renewable Sustainable Energy Rev.*, 2018, **97**, 259–275.
- 14 S. He and J. Cai, *Ann. Nucl. Energy*, 2020, **140**, 107303.
- 15 *The Rare Earth Elements: Fundamentals and Applications*, Ed. D. A. Atwood, John Wiley & Sons, Ltd, Chichester, 2012.
- 16 *The Chemistry of the Actinide and Transactinide Elements*, Ed. L. R. Morss, N. M. Edelstein and J. Fuger, Springer, Dordrecht, 3rd edn, 2006.
- 17 C. Aitken, J. F. Harrod and E. Samuel, *J. Organomet. Chem.*, 1985, **279**, C11–C13.
- 18 C. R. Groom, I. J. Bruno, M. P. Lightfoot and S. C. Ward, *Acta Crystallogr., Sect. B: Struct. Sci., Cryst. Eng. Mater.*, 2016, **72**, 171–179. Search date 13th March 2023.
- 19 B. L. L. Réant, V. E. J. Berryman, J. A. Seed, A. R. Basford, A. Formanuik, A. J. Wooles, N. Kaltsoyannis, S. T. Liddle and D. P. Mills, *Chem. Commun.*, 2020, **56**, 12620–12623.
- 20 G. K. Gransbury, B. L. L. Réant, A. J. Wooles, J. Emerson-King, N. F. Chilton, S. T. Liddle and D. P. Mills, *Chem. Sci.*, 2023, **14**, 621–634.
- 21 T. Imori, R. H. Heyn, T. D. Tilley and A. L. Rheingold, *J. Organomet. Chem.*, 1995, **493**, 83–89.
- 22 M. F. Lappert, C. J. Pickett, P. I. Riley and P. I. W. Yarrow, *J. Chem. Soc., Dalton Trans.*, 1981, 805–813.
- 23 A. Antiñolo, M. F. Lappert, A. Singh, D. J. Winterborn, L. M. Engelhardt, C. L. Raston, A. H. White, A. J. Carty and N. J. Taylor, *J. Chem. Soc., Dalton Trans.*, 1987, 1463–1472.
- 24 P. C. Blake, M. F. Lappert, R. G. Taylor, J. L. Atwood, W. E. Hunter and H. Zhang, *J. Chem. Soc., Dalton Trans.*, 1995, 3335–3341.
- 25 C. Marschner, *Eur. J. Inorg. Chem.*, 1998, 221–226.
- 26 C. H. Winter, X.-X. Zhou, A. A. Dobbs and M. J. Heeg, *Organometallics*, 1991, **10**, 210–214.
- 27 C. J. Windorff, M. R. MacDonald, J. W. Ziller and W. J. Evans, *Z. Anorg. Allg. Chem.*, 2017, **643**, 2011–2018.
- 28 M. Schlosser and J. Hartmann, *Angew. Chem., Int. Ed. Engl.*, 1973, **12**, 508–509.
- 29 I. Krossing, H. Brands, R. Feuerhake and S. Koenig, *J. Fluor. Chem.*, 2001, **112**, 83–90.
- 30 J. B. Lambert, S. Zhang and S. M. Ciro, *Organometallics*, 1994, **13**, 2430–2443.
- 31 J. K. Peterson, M. R. MacDonald, J. W. Ziller and W. J. Evans, *Organometallics*, 2013, **32**, 2625–2631.
- 32 D. E. Bergbreiter and J. M. Killough, *J. Am. Chem. Soc.*, 1978, **100**, 2126–2134.
- 33 P. C. Blake, M. F. Lappert, J. L. Atwood and H. Zhang, *J. Chem. Soc., Chem. Commun.*, 1986, 1148–1149.
- 34 R. R. Langeslay, M. E. Fieser, J. W. Ziller, F. Furche and W. J. Evans, *Chem. Sci.*, 2015, **6**, 517–521.
- 35 G. J. Pindado, M. Thornton-Pett and M. Bochmann, *J. Chem. Soc., Dalton Trans.*, 1997, 3115–3128.
- 36 A. Pastor, A. F. Kiely, L. M. Henling, M. W. Day and J. E. Bercaw, *J. Organomet. Chem.*, 1997, **528**, 65–75.
- 37 F. H. Elsner, T. D. Tilley, A. L. Rheingold and S. J. Geib, *J. Organomet. Chem.*, 1988, **358**, 169–183.
- 38 H.-G. Woo, R. H. Heyn and T. D. Tilley, *J. Am. Chem. Soc.*, 1992, **114**, 5698–5707.
- 39 B. K. Campion, J. Falk and T. D. Tilley, *J. Am. Chem. Soc.*, 1987, **109**, 2049–2056.
- 40 C. Kayser, D. Frank, J. Baumgartner and C. Marschner, *J. Organomet. Chem.*, 2003, **667**, 149–153.
- 41 P. Pykkö and M. Atsumi, *Chem. – Eur. J.*, 2009, **15**, 186–197.
- 42 K. C. Jantunen, B. L. Scott and J. L. Kiplinger, *J. Alloys Compd.*, 2007, **444–445**, 363–368.
- 43 B. L. L. Réant, V. E. J. Berryman, A. R. Basford, L. E. Nodaraki, A. J. Wooles, F. Tuna, N. Kaltsoyannis, D. P. Mills and S. T. Liddle, *J. Am. Chem. Soc.*, 2021, **143**, 9813–9824.
- 44 B. L. L. Réant, A. J. Wooles, S. T. Liddle and D. P. Mills, *Inorg. Chem.*, 2023, **62**, 137–146.
- 45 M. Kaupp, O. L. Malkina, V. G. Malkin and P. Pykkö, *Chem. – Eur. J.*, 1998, **4**, 118–126.
- 46 *MestReNova v.14.3.1*, Mestrelab Research, S. L. Santiago de Compostela, Spain, 2022; <https://mestrelab.com/software/mnova/>.
- 47 P. B. Hitchcock, M. F. Lappert, L. Maron and A. V. Protchenko, *Angew. Chem., Int. Ed.*, 2008, **47**, 1488–1491.
- 48 *CrysAlis PRO*, Agilent Technologies Ltd, Yarnton, Oxfordshire, England, 2014.
- 49 G. M. Sheldrick, *Acta Crystallogr., Sect. A: Found. Adv.*, 2015, **71**, 3–8.
- 50 G. M. Sheldrick, *Acta Crystallogr., Sect. C: Struct. Chem.*, 2015, **71**, 3–8.
- 51 O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. K. Howard and H. Puschmann, *J. Appl. Crystallogr.*, 2009, **42**, 339–341.
- 52 L. J. Farrugia, *J. Appl. Crystallogr.*, 2012, **45**, 849–854.



- 53 Persistence of Vision Raytracer Pty. Ltd, *Persistence of Vision Raytracer*, v.3.7, Williamstown, VIC, Australia, 2013. Retrieved from <https://www.povray.org/download/>.
- 54 C. Fonseca Guerra, J. G. Snijders, G. te Velde and E. J. Baerends, *Theor. Chem. Acc.*, 1998, **99**, 391–403.
- 55 G. te Velde, F. M. Bickelhaupt, E. J. Baerends, A. C. Fonseca Guerra, S. J. A. van Gisbergen, J. G. Snijders and T. Ziegler, *J. Comput. Chem.*, 2001, **22**, 931–967.
- 56 E. van Lenthe, E. J. Baerends and J. G. Snijders, *J. Chem. Phys.*, 1993, **99**, 4597.
- 57 E. van Lenthe, E. J. Baerends and J. G. Snijders, *J. Chem. Phys.*, 1994, **101**, 9783.
- 58 E. van Lenthe, A. E. Ehlers and E. J. Baerends, *J. Chem. Phys.*, 1999, **110**, 8943.
- 59 S. H. Vosko, L. Wilk and M. Nusair, *Can. J. Phys.*, 1980, **58**, 1200–1211.
- 60 A. D. Becke, *Phys. Rev. A*, 1988, **38**, 3098.
- 61 J. P. Perdew, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1986, **33**, 8822.
- 62 E. D. Glendening, J. K. Badenhoop, A. E. Reed, J. E. Carpenter, J. A. Bohmann, C. M. Morales, C. R. Landis and F. Weinhold, *NBO v.6.0*, Theoretical Chemistry Institute, University of Wisconsin, Madison, WI, USA, 2013; <https://nbo6.chem.wisc.edu/>.
- 63 R. F. W. Bader, *Atoms in Molecules: A Quantum Theory*, Oxford University Press, New York, 1990.
- 64 R. F. W. Bader, *J. Phys. Chem. A*, 1998, **102**, 7314–7323.

