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A highly Li⁺-conductive HfNb₂₄O₆₂ anode material for superior Li⁺ storage†

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Highly Li⁺-conductive HfNb₂₄O₆₂ is explored as a new intercalation-type niobium-based oxide anode material for superior Li⁺ storage. HfNb₂₄O₆₂ owns a Wadsley-Roth shear structure with a large unit-cell volume, leading to a large Li⁺ diffusion coefficient. HfNb₂₄O₆₂ shows a large capacity, safe operating potential, high rate performance and good cyclability.

Safe lithium-ion batteries (LIBs) with high power and energy density for electric vehicles (EVs) and hybrid electric vehicles (HEVs) have attracted intense attention. However, the traditional graphite anode material is limited in such applications due to its low kinetics and unsafe lithiation potential.² As a promising substitute for graphite, $\text{Li}_4\text{Ti}_5\text{O}_{12}$ is appealing due to its safe operating potential ($\sim 1.57 \text{ V}$) and high rate performance after proper modifications.³ However, the small theoretical capacity of Li₄Ti₅O₁₂ (175 mA h g⁻¹) has limited its large-scale applications. Therefore, new anode materials for safe, high-power and high-energy LIBs have been pursued.

In recent years, intercalation-type niobium-based oxides, such as $Ti_2Nb_{2x}O_{4+5x}$ (x = 2, 5 and 24), 4-11 MNb₁₁O₂₉ (M = Al, Cr, Fe and Ga), 12-15 and W₅Nb₁₆O₅₅, 16 have been developed as potential candidates to replace Li₄Ti₅O₁₂ due to their comparably safe operating potentials and much larger theoretical/practical capacities. Niobium is an appealing redox center in electrode materials because of its safe potential window in the range of 1.0-2.0 V for the redox couples of Nb4+/Nb5+ and Nb3+/Nb4+, which can avoid the formation of dangerous lithium dendrites.4 In light of multivalent properties, niobium-based oxides can deliver large theoretical capacities of 374-416 mA h g^{-1} , which are close to that of graphite and significantly surpass that of Li₄Ti₅O₁₂. In addition, with reported open Wadsley-Roth shear crystal structures consisting of 96-100% MO₆ octahedra and 0-4% MO₄ tetrahedra, niobium-based oxides have been identified as typical intercalation-pseudocapacitive materials, which are beneficial for the capacity and fast-charging property. Despite such attractive features, niobium-based oxide anodes still show poor rate performance due to their intrinsically low Li⁺ conductivities. Therefore, developing new niobium-based oxide anode materials with high Li⁺ conductivities, Li⁺-storage capacities, rate performance, and long-term stability is essential.

In this study, HfNb24O62 was synthesized by a facile and costeffective solid-state reaction, and applied as a novel niobiumbased oxide anode material with a high Li⁺ conductivity to boost the Li⁺ storage. In octahedral coordination, the Hf⁴⁺ ionic radius (0.710 Å) is larger than those of Ti^{4+} (0.605 Å), Al^{3+} (0.535 Å), Cr^{3+} (0.615 Å), Fe³⁺ (0.550 Å), Ga³⁺ (0.620 Å) and W⁶⁺ (0.600), ¹⁷ leading to increased lattice constants and unit-cell volume in HfNb24O62. A larger unit-cell volume usually corresponds to faster Li⁺ diffusivity, benefiting the Li⁺ transport kinetics. ¹⁸ Moreover, HfNb₂₄O₆₂ exhibits a monoclinic Wadsley-Roth shear crystal structure (space group of C2) consisting of 96% edge-sharing (Hf,Nb)O₆ octahedra linked by 4% (Hf,Nb)O₄ tetrahedra, in which Hf⁴⁺ and Nb⁵⁺ are disorder located. This three-dimensional skeleton bestows fast Li⁺ diffusivity and high structural stability, guaranteeing fast lithiation/delithiation and excellent cyclability. Both the HfNb₂₄O₆₂/Li half-cell and LiNi_{0.5}Mn_{1.5}O₄//HfNb₂₄O₆₂ full-cell exhibit superior electrochemical properties for fast-charging, large-capacity, safe and durable Li⁺ storage.

The detailed crystal structure of the as-prepared HfNb₂₄O₆₂ material was studied by XRD and a Rietveld refinement of the XRD pattern (Fig. 1a).¹⁹ The characteristic peaks of HfNb₂₄O₆₂

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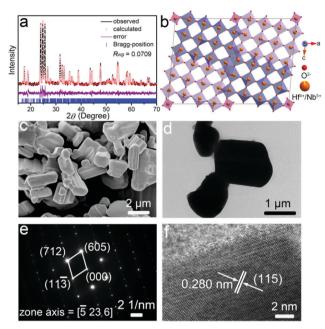


Fig. 1 (a) XRD pattern, (b) crystal structure, (c) FESEM image, (d) TEM image, (e) SAED pattern, and (f) HRTEM image of HfNb₂₄O₆₂. R_{wp}: weighted profile residual.

are consistent with those of TiNb₂₄O₆₂ (space group of C2).⁶ HfNb₂₄O₆₂ owns a robust crystal structure, with the structural unit having a 3 × 4 (Hf,Nb)O₆ octahedron-block and a half (Hf,Nb)O₄ tetrahedron at the block corner (Fig. 1b). The lattice parameters were Rietveld-refined to be a = 29.92508(125) Å, $b = 3.82525(14) \text{ Å}, c = 21.21133(87) \text{ Å}, \beta = 95.068(5)^{\circ}, \text{ and}$ $V = 2418.588(167) \text{ Å}^3$. Detailed lattice parameters are given in Table S1 (ESI†). Due to the larger Hf⁴⁺ (0.710 Å) ionic radius than Ti^{4+} (0.605 Å), the unit-cell volume of $HfNb_{24}O_{62}$ is larger than that of TiNb₂₄O₆₂, which implies wider Li⁺ transport pathways in the HfNb₂₄O₆₂ lattice. The survey XPS spectrum of HfNb₂₄O₆₂ reveals the presence of Hf, Nb, O and C (reference) elements (Fig. S1a, ESI†). The detailed Hf-4f (Fig. S1b, ESI†) and Nb-3d (Fig. S1c, ESI†) spectra indicate that the Hf and Nb valence states are +4 and +5,4,20 respectively.

The FESEM (Fig. 1c) and TEM (Fig. 1d) images reveal that the HfNb₂₄O₆₂ material is composed of irregular and dense particles with particle sizes varying from ~ 2 to $\sim 10 \mu m$, giving a low specific surface area of 0.38 m² g⁻¹ (Fig. S2, ESI†). The corresponding SAED pattern in Fig. 1e confirms high crystallinity and reveals a single-crystalline feature of the HfNb24O62 primary particles. An interplanar spacing of 0.280 nm measured from the HRTEM image (Fig. 1f) corresponds to the (115) crystallographic plane of HfNb₂₄O₆₂. Besides, the EDX mappings (Fig. S3, ESI†) indicate that the Nb, O and Hf elements are welldispersed in the whole microparticle, verifying the formation of pure HfNb₂₄O₆₂.

The electrochemical properties of HfNb₂₄O₆₂ were mainly studied using half-cells. Fig. 2a shows the CV curves of the HfNb₂₄O₆₂/Li cell for the initial four cycles at a sweep rate of $0.2~{\rm mV~s^{-1}}$. The difference in the discharge portion of the first CV curve from the subsequent cycles is likely due to the

incomplete Li⁺ extraction from the HfNb₂₄O₆₂ lattice in the first cycle.21 After the first cycle, the peak positions and corresponding currents are very stable, suggesting good electrochemical reversibility and cyclability of HfNb₂₄O₆₂. In the subsequent sweeps, two highly overlapping CV peak pairs located at $\sim 2.0/\sim 2.0$ and 1.77/ 1.55 V were assigned to the reversible transformation for the redox couple of Nb4+/Nb5+. The shoulder peaks located at 1.34/1.14 V correspond to the redox couple of Nb³⁺/Nb⁴⁺. ¹² Consequently, the estimated HfNb₂₄O₆₂ mean operating potential (~1.66 V) based on the intermediate potential of two intensive peaks at 1.77/1.55 V is close to those of $\text{Li}_4\text{Ti}_5\text{O}_{12} \ (\sim 1.57 \ \text{V})^3 \ \text{TiNb}_2\text{O}_7 \ (\sim 1.64 \ \text{V})^5$ $TiNb_{24}O_{62}$ (~1.66 V)⁶ and $Ti_2Nb_{10}O_{29}$ (~1.70 V). Such a reasonably high operating potential of HfNb₂₄O₆₂ indicates its high safety performance.

Fig. 2b displays the initial three-cycle discharging/charging curves of the HfNb24O62/Li cell recorded at 0.1C. The observed typical potential plateau at $\sim 1.66 \text{ V}$ in the discharging/charging curves resulted from a double-phase transformation reaction. The sloping regions at 3.0-1.67 and 1.65-0.8 V were assigned to two different solid-solution reactions.8 HfNb24O62 shows an initial Coulombic efficiency of 93.8% and large reversible capacity of 272 mA h g^{-1} at 0.1C. This practical capacity of HfNb₂₄O₆₂ is >100 mA h g⁻¹ larger than that of Li₄Ti₅O₁₂ and surpasses the majority of the reported intercalation-type niobium-based oxide anode materials (Fig. 2c). 3,4,6,9,10,14-16,22,23 The high initial Coulombic efficiency can be due to the fact that little SEI layers formed on the HfNb₂₄O₆₂ particle surfaces above 0.8 V.

The rate performance of HfNb24O62 was investigated at various current rates for every ten cycles, as shown in Fig. 2d and e. HfNb₂₄O₆₂ delivers large reversible capacities of 223, 195, 174, 138 and 105 mA h g⁻¹ at 0.5, 1, 2, 5 and 10C, respectively. When the current rate gradually returns from 10C to 0.5C, the capacity is recovered to its original value. When cycled at 1C, HfNb₂₄O₆₂ shows a small capacity change from 193 to 182 mA h g^{-1} after 100 cycles, giving 94.3% capacity retention (Fig. 2f and Fig. S4, ESI†). When cycled at 10C, HfNb₂₄O₆₂ shows 87.1% capacity retention even after 500 cycles during the prolonged cycling (Fig. 2g and Fig. S5, ESI†). Such superior cyclability can be due to the stable crystal structure and good Li⁺-transport kinetics of HfNb₂₄O₆₂.

To interpret the high rate performance of HfNb₂₄O₆₂, its Li⁺-storage kinetics was analyzed by additional CV experiments recorded at different sweep rates (Fig. 2h). It was found that the cathodic/anodic CV peak current exhibited power law dependence on the sweep rate with an exponent of 0.67/0.74 (Fig. 2i),7,24 suggesting that the pseudocapacitive behaviour significantly contributed to the fast charge storage of HfNb24O62 since it is well known that exponents of 0.5 and 1 indicate diffusion-controlled and pseudocapacitive charge storage, respectively.

A galvanostatic intermittent titration technique (GITT) with a current pulse at 0.1C for 10 min between rest intervals for 20 min was applied to elucidate the Li⁺ diffusion coefficient $(D_{\rm Li})$ of HfNb₂₄O₆₂ during the initial two cycles. Based on Fick's second law (see ESI†), 25 the evolving D_{Li} of HfNb₂₄O₆₂ at each potential in the discharging (lithiation) and charging (delithiation) processes is determined and plotted in Fig. 2j. The D_{Li} values show

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a0.6 Nb⁴⁺/Nb 2nd €^{2.5} Potential (Nb⁴⁺/Nb⁵ J-0.2 2nd -0.4 3rd _{-0.6}]0.2 mV s 4th 50 100 150 200 Specific capacity (mAh g 1) 150 Specific capacity (mAh g **e** 300 Specific capacity (mAh g⁻¹) 0.5C 0.5C 1C 1C W₁₅₀ 2C 2C 5C charge capacity discharge capacity 🧖 10C 100 Specific 50 charge capacity discharge capacity Cycle numbers 80 80 40 50 Cycle numbers **g** 150 h -2.70 • anodic Specific capacity (mAh g⁻¹) cathodic Æ Current (mA) charge capacity 87 19 discharge capacity (peak ---- 0.4 mV s go 0 cycled at 10C 1.1 mV s 200 300 Cycle numbers Potential (V) -0.6 log (sweep rate, mV s⁻¹) 2nd lithiation 1st lithiation 2nd delithiation 1st delithiation Potential (V) $D_{\rm Li^+} (10^{-12} {\rm cm^2 s^{-1}})$

Z0000 Time (min) Fig. 2 Electrochemical characterizations of the HfNb₂₄O₆₂/Li cell: (a) CVs at 0.2 mV s⁻¹, (b) discharging/charging curves at 0.1C, (c) comparisons of reversible capacity of HfNb24O62 with other intercalation-type oxide anode materials, (d) discharging/charging curves at different current rates, (e) rate performance, cyclability at (f) 1C and (g) 10C, (h) CVs at different sweeping rates, (i) exponential relationship between CV peak current and sweep rate, and (j) GITT curves and corresponding Li⁺ diffusion coefficients (initial two cycles).

1500

a similar variation trend in the two cycles, and minor values appear in the plateau region ($\sim 1.7 \text{ V}$), where the Li⁺ interaction with the host matrix is strong during the two-phase transition reaction. During the first lithiation/delithiation processes, the D_{Li} value varies from 1.51 \times 10⁻¹³ to 2.91 \times 10⁻¹² cm² s⁻¹, averaging at $1.61 \times 10^{-12} \text{ cm}^2 \text{ s}^{-1}$. The average D_{Li} value in the second lithiation/delithiation processes (1.70 \times 10⁻¹² cm² s⁻¹) is slightly larger than that of the initial cycle, but the curves are similar to those in the initial cycle. It is noteworthy that the Li⁺ diffusion coefficient of HfNb24O62 is significantly larger than those of the reported niobium-based oxide anode materials (Table S2, ESI†), which can be ascribed to its open Wadsley-Roth shear crystal structure with an enlarged unit-cell volume. Clearly, the fast Li⁺ diffusivity together with the significant pseudocapacitive behaviour of HfNb₂₄O₆₂ greatly contributes to its superior Li⁺ storage (especially its rate performance).

To demonstrate the practical application of HfNb₂₄O₆₂, a full cell with a HfNb₂₄O₆₂ anode and a LiNi_{0.5}Mn_{1.5}O₄ cathode was fabricated (Fig. 3a). The LiNi_{0.5}Mn_{1.5}O₄//HfNb₂₄O₆₂ full cell affords a large charge capacity of 213 mA h g⁻¹ with an average operating voltage of \sim 3.0 V at 0.1C (Fig. 3b). This high voltage agrees with the CV result (Fig. 3c). As the current rate gradually increases to 0.5, 1, 2, 5 and even 10C, the reversible capacity retains 143, 130, 116, 95 and 78 mA h g^{-1} , respectively (Fig. 3d). When the current rate returns from 10C to 0.5C, the obtained capacity of 141 mA h g^{-1} indicates good electrochemical reversibility (Fig. 3e). The full cell also presents outstanding cyclability at both 1C (83.4% capacity retention after 100 cycles, Fig. 3f) and 5C (84.7% capacity retention after 500 cycles, Fig. 3g). The full cell can power a green light-emitting diode (LED) after being cycled 500 times at 5C (Fig. 3g inset).

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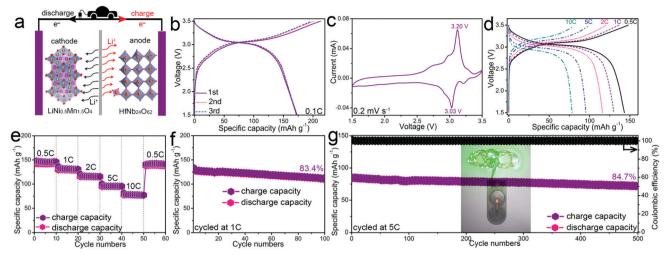
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In summary, highly-Li⁺-conductive HfNb₂₄O₆₂ is demonstrated as a novel anode material to realize superior Li⁺ storage.

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 $\textbf{Fig. 3} \hspace{0.2cm} \textbf{(a)} \hspace{0.2cm} \hspace{$ (b) initial three-cycle charging/discharging curves at 0.1C, (c) CV at 0.2 mV s⁻¹, (d) charging/discharging curves at different current rates, (e) rate performance, and cyclability at (f) 1C and (g) 5C (inset: green LED lit up by the full cell).

Both the open Wadsley-Roth shear crystal structure and enlarged unit-cell volume lead to favorable Li⁺ conduction. Benefiting from the robust host framework and fast ion diffusion pathways, HfNb₂₄O₆₂ exhibits prominent electrochemical properties. At 0.1C, it displays a large capacity of 272 mA h g⁻¹ with a relatively safe operating potential of ~ 1.66 V and a high initial Coulombic efficiency of 93.8%. Even at 10C, a remarkable reversible capacity of 105 mA h g⁻¹ is preserved. Furthermore, it exhibits good long-term cyclability, as indicated by 87.1% capacity retention after 500 cycles at 10C. More importantly, a $\text{LiNi}_{0.5}\text{Mn}_{1.5}\text{O}_4\text{//HfNb}_{24}\text{O}_{62}$ full cell also exhibits high rate performance with 213 mA h g^{-1} at 0.1C and 78 mA h g^{-1} at 10C, and good long-term cyclability with 84.7% capacity retention after 500 cycles at 5C. This work can benefit the future designs of highly-Li⁺-conductive and fast-charging electrode

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Conflicts of interest

There are no conflicts to declare.

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