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# Photodissociation of permanganate ( $\text{MnO}_4^-$ ) produces the manganese dioxide anion ( $\text{MnO}_2^-$ ) in an excited triplet state†

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Photoelectron imaging, electron action spectroscopy and electronic structure calculations are used to probe the structure and dynamics of  $\text{MnO}_4^-$ . Following excitation to the first bright absorption band of  $\text{MnO}_4^-$  ( $1^1\text{T}_2$ ), photodetachment, *via* ground state electron loss, and photodissociation, to produce  $\text{MnO}_2^-$ , are both observed to occur simultaneously.  $\text{MnO}_2^-$  is produced in an excited electronic state, identified as a triplet state, which indicates that the dissociation proceeds on singlet potential energy surfaces *via* spin conservation. Furthermore, electronic structure calculations indicate that both photodetachment and photodissociation are multiple photon processes that are mediated by the same  $1^1\text{T}_2$  excited state. Taken together this data indicates that photodissociation of  $\text{MnO}_4^-$  occurs *via* a statistical dissociation on the  $\text{MnO}_4^-$  ground state at visible wavelengths.

## Introduction

Permanganate ( $\text{MnO}_4^-$ ) is an archetypal transition metal complex, familiar to chemistry students worldwide for its deep purple colour. The well-known absorption spectrum of  $\text{KMnO}_4(\text{aq})$  shows a strong visible absorption band near  $h\nu = 2.25$  eV ( $\lambda \sim 550$  nm), assigned to the first fully allowed transition from the tetrahedral ground state of  $\text{MnO}_4^-$ ,  $1\text{A}_1$ , to the  $1^1\text{T}_2$  excited state.<sup>1</sup> The  $1\text{A}_1 \rightarrow 1^1\text{T}_2$  absorption band displays vibronic structure predominantly associated with a vibrational progression of the Mn–O stretching mode ( $770\text{ cm}^{-1}$ ).<sup>1</sup> Photoexcitation to the  $1^1\text{T}_2$  state has also been known to lead to  $\text{MnO}_2^- + \text{O}_2$  in aqueous solution for many years, although the underlying photodissociation mechanism remains unclear.<sup>2,3</sup> Early work postulated that dissociation occurred on the ground state *via* a unimolecular decomposition after the rapid formation of a vibrationally excited ground state anion, due to an increasing propensity for  $\text{MnO}_2^-$  formation with increasing photon energy.<sup>3,4</sup> This mechanism has been challenged, as the quantum yield for photodissociation seemed linked to the specific electronically excited state accessed, leading to speculation that the red-tail of a higher-lying UV transition was responsible for dissociation in the visible band.<sup>5</sup> It has also been postulated that electron loss may precede dissociation.<sup>5</sup> More recent work has proposed that  $\text{MnO}_2^-$  production occurs following the formation of a long-lived intermediate, potentially a Mn(v) peroxo-complex, which undergoes

reductive elimination.<sup>4,6</sup> However, this long-lived intermediate has only been observed following excitation at UV wavelengths suggesting this process might not contribute in the visible region.<sup>4,6</sup> Recent femtosecond transient absorption spectroscopy showed that excitation to the  $1^1\text{T}_2$  state is followed by a rapid ( $\sim 16$  ps) return to the ground state, *via* a cascade of radiationless transitions.<sup>4</sup> Supported by quantum chemical calculations, the mechanism was determined to involve very fast internal conversion (IC) of the  $1^1\text{T}_2$  state to a lower-lying  $1^1\text{T}_1$  state, which is accompanied by a Jahn–Teller (JT) distortion. Subsequently, this  $1^1\text{T}_1$  state undergoes intersystem crossing (ISC) to the triplet-manifold, followed by IC to the lowest lying triplet state ( $3^1\text{T}_1$ ), which is more stable than the lowest energy singlet in the JT-distorted geometry. Afterwards, the JT-distorted triplet decays by ISC back to the  $1\text{A}_1$  electronic ground state in  $\sim 16$  ps with the accompanying structural relaxation to the tetrahedral geometry.<sup>4</sup> These dynamics are summarised in Fig. 1.

In order to elucidate the intrinsic photodissociation mechanism of permanganate, a number of experimental and computational studies of isolated (gas phase)  $\text{MnO}_4^-$  have also been performed.<sup>7–9</sup> Fragment ion action spectroscopy by Houmoller *et al.* showed that photodissociation of  $\text{MnO}_4^-$  produces both  $\text{MnO}_3^-$  and  $\text{MnO}_2^-$  following  $1\text{A}_1 \rightarrow 1^1\text{T}_2$  excitation ( $2.48 \leq h\nu \leq 2.07$  eV,  $500 \leq \lambda \leq 600$  nm).<sup>7,10</sup> Fluence dependent measurements indicated that the photodissociation is a multiple photon process in this wavelength region.<sup>7,8</sup> This was in agreement with theoretical studies which suggest that both the  $\text{MnO}_4^-$  dissociation asymptotes,  $\text{MnO}_2^- + \text{O}_2$  and  $\text{MnO}_3^- + \text{O}$ , lie approximately 5 eV above the  $\text{MnO}_4^-$  ground state.<sup>8</sup>

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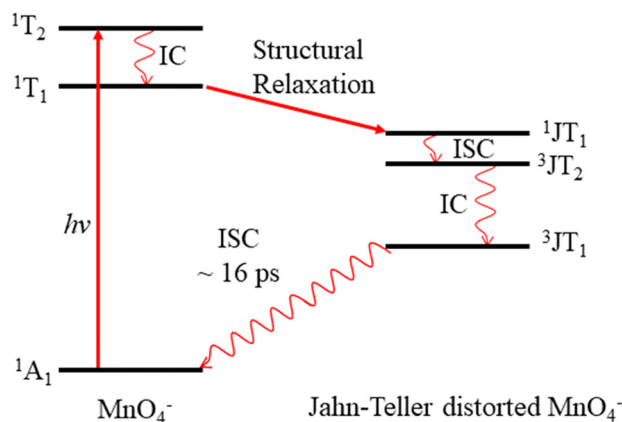


Fig. 1 Excitation of  $\text{MnO}_4^-$  in the visible to  $^1T_2$ , followed by rapid relaxation via the  $^3JT_1$  state of a Jahn–Teller distorted  $\text{MnO}_4^-$  back to  $^1A_1$ , via structural relaxation, internal conversion (IC) and intersystem crossing (ISC). The transition from  $^1A_1$  to  $^1T_1$  is symmetry forbidden.

Photoelectron spectroscopy of  $\text{MnO}_4^-$  has also been performed by Gutsev *et al.*, reporting the electron affinity of  $\text{MnO}_4^-$  to be  $4.8 \pm 0.1$  eV, as well as calculations which suggested that photodissociation of  $\text{MnO}_4^-$  in the visible must involve a spin-flip and be a multiple photon process.<sup>8,11,12</sup> Finally, Gutsev *et al.* also measured the photoelectron spectrum of  $\text{MnO}_4^-$ , where they were able to discern that  $\text{MnO}_2^-$  was produced in both the ground electronic state ( $^5B_2$ ) and the first excited triplet state ( $^3B_1$ ), in a ratio which depends upon the initial temperature of the ions.<sup>11</sup>

Despite the wealth of studies investigating the structure and dynamics of  $\text{MnO}_4^-$ , in both the solution and gas phase, fundamental questions remain about the photodissociation process. For example, what is the dissociation mechanism following excitation at visible wavelengths in the  $^1A_1 \rightarrow ^1T_2$  band and what is the quantum state distribution of the products. Here, using a combination of photoelectron imaging, electronic structure calculations and electron action spectroscopy on  $\text{MnO}_4^-$ , we show that photodissociation forms  $\text{MnO}_2^-$  in an electronically excited state, which is consistent with a thermal unimolecular dissociation mechanism on the  $^1A_1$  ground state.

## Experimental methods

The experimental apparatus has been described in detail elsewhere and will only be summarised here.<sup>13,14</sup> Anions are produced *via* electrospray ionization of a 5 mM solution of  $\text{KMnO}_4$  in acetonitrile. Anions enter the apparatus *via* a capillary, before being guided and trapped using a series of RF guides and subsequent acceleration to 2.2 kV using a Wiley–McLaren time of flight spectrometer.<sup>15</sup> Mass selected anions are intersected with the variable wavelength output of a nano-second optical parametric oscillator (OPO) pumped *via* the third harmonic of a Nd:YAG, operating at 10 Hz. Ejected electrons are focused onto the position sensitive microchannel

plate and phosphor screen detector using a velocity map imaging setup. From the resulting photoelectron images, the electron kinetic energy (eKE) and photoelectron angular distribution (PAD) are extracted using the polar onion peeling algorithm.<sup>16</sup> From the eKE spectrum and the photon energy,  $h\nu$ , the electron binding energy (eBE) can be extracted ( $\text{eKE} + \text{eBE} = h\nu$ ). The PADs are characterised by an anisotropy parameter ( $-1 \leq \beta_2 \leq 2$ ) and provide information on the molecular orbital from which the ejected photoelectron originated.<sup>17</sup> Photoelectron imaging is performed over a range of photon energies, to determine how the spectral features shift with wavelength. As the features in the spectra presented here are consistent across the wavelengths, only a sample of the spectra will be reported.

The electron action spectrum of  $\text{MnO}_4^-$  was recorded by monitoring the yield of photoelectrons, with the electron imaging detector, as a function of wavelength. The data was recorded whilst scanning the OPO at  $0.1 \text{ nm s}^{-1}$  and is smoothed by a 20-point moving average. The laser power from the OPO will vary somewhat across the wavelengths studied and is independently measured using a power meter.

Electronic structure calculations to determine the relative energetics of the possible electron loss and photodissociation channels of  $\text{MnO}_4^-$  were performed using DFT at the B3LYP level of theory with the aug-cc-pVTZ basis set *via* Gaussian 16.<sup>18–20</sup> For each species the geometry was optimised and confirmed to be a true minima by vibrational analysis. All calculated energies are zero point energy corrected. This theoretical methodology has previously been used to study the electronic and nuclear structure of  $\text{MnO}_3^-$  and  $\text{MnO}_3$ , where it was found to offer reliable results.<sup>21</sup>

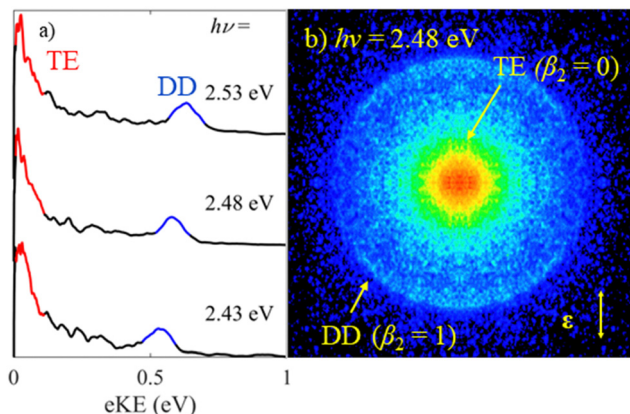
## Results

### (a) Photoelectron imaging of $\text{MnO}_4^-$

Fig. 2 shows photoelectron spectra of  $\text{MnO}_4^-$  as well as an exemplar raw photoelectron image. Despite the large electron affinity of  $\text{MnO}_4^-$ ,<sup>12</sup> photoelectrons are clearly ejected from  $\text{MnO}_4^-$  after irradiation with  $h\nu \sim 2.5$  eV ( $\lambda \sim 500$  nm) light. There are two dominant features in the photoelectron spectra shown in Fig. 2(a): a broad band of electrons peaking near  $\text{eKE} \sim 0$  eV; and a relatively narrow feature at fixed electron binding energy,  $\text{eBE} = 1.9 \pm 0.1$  eV (corresponding to an  $\text{eKE}$  of 0.6 eV at  $h\nu = 2.48$  eV). From the image in Fig. 2(b), the feature near  $\text{eKE} \sim 0$  eV and labelled thermionic emission is isotropic ( $\beta_2 = 0$ ), while the outer ring at higher  $\text{eKE}$  ( $\text{eBE} = 1.9$  eV), labelled direct detachment, is anisotropic and shows that photoelectrons are emitted predominantly parallel to the polarisation axis of the light ( $\beta_2 \approx +1.0$ ).

The spectral feature peaking near  $\text{eKE} \sim 0$  eV has a characteristic distribution (both in terms of energy and PAD) associated with unimolecular loss of electrons from a vibrationally hot ground state (thermionic emission).<sup>22–25</sup> The second narrow spectral feature increases in  $\text{eKE}$  with increasing  $h\nu$ , indicating a direct photodetachment process. Given that



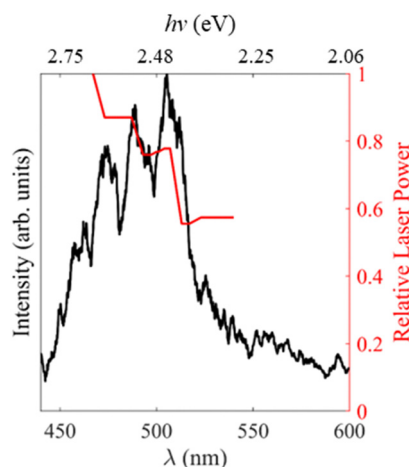


**Fig. 2** (a) Photoelectron spectrum of  $\text{MnO}_4^-$  recorded at  $h\nu = 2.53$  eV (490 nm), 2.48 eV (500 nm) and 2.43 eV (510 nm). (b) Raw symmetrised photoelectron image of  $\text{MnO}_4^-$  recorded at  $h\nu = 2.48$  eV, where the laser polarisation axis is indicated by the double arrow. The spectral features associated with thermionic emission (TE) and direct detachment (DD) are highlighted in red and blue respectively, in both the photoelectron spectra and image.

eBE = 1.9 eV is significantly lower than the reported electron affinity of  $\text{MnO}_4$ ,<sup>12</sup> this feature is attributable to a multiple-photon process in which a final photon detaches an electron either from an excited state of  $\text{MnO}_4^-$  with eBE of 1.9 eV or from a photofragment with this eBE. Photodissociation of  $\text{MnO}_4^-$  produces  $\text{MnO}_2^-$ , making this a likely candidate for an anionic fragment.<sup>3,4,6,7</sup> It is possible that some, or all, of the thermionic emission arises from this fragment too.

### (b) Electron action spectroscopy

In order to gain insights into the electronic states which mediate the electron loss pathways observed in the photoelectron spectra of  $\text{MnO}_4^-$ , we recorded the electron action spectrum, *i.e.*, the yield of photoelectrons as a function of wavelength, which is shown in Fig. 3. A gradual increase in



**Fig. 3** Electron action spectrum of  $\text{MnO}_4^-$  recorded between  $h\nu = 2.81 - 2.06$  eV ( $\lambda = 440$  and 600 nm) in black. The relative laser power is shown in red on the right.

electron yield is seen with increasing  $h\nu$ , followed by a large increase around  $h\nu = 2.4$  eV ( $\lambda = 520$  nm), beyond which vibrational structure becomes apparent, with a spacing of  $\sim 770$   $\text{cm}^{-1}$ . Both photo-processes observed in the photoelectron spectra in Fig. 2 (*i.e.*, thermionic emission and multiple-photon photodetachment) contribute electrons to the electron action spectrum. The laser power variation across the spectrum is shown in red on the right-side axis, but the spectrum is not normalised to account for this, due to the contribution of multiple photon processes.

### (c) Calculations

The calculated relative energetics for the photodetachment and photodissociation channels of  $\text{MnO}_4^-$  are shown in Table 1 and compared to experimental data where available.<sup>8,11,12</sup> In order to test the validity of our methodology, we computed the bond dissociation energy of  $\text{O}_2$  and found it to be  $D_0 = 5.2$  eV, which is close to the experimental value of 5.12 eV.<sup>26</sup> There is also good agreement between our calculated electron affinities and the experimentally observed detachment energies for  $\text{MnO}_2^-$ ,  $\text{MnO}_3^-$  and  $\text{MnO}_4^-$ .<sup>11,12</sup> Finally, we have previously shown that the computational methodology predicts electronic and geometric structures with good accuracy for  $\text{MnO}_3$  and its anion.<sup>21</sup> The high electron affinity of  $\text{MnO}_4$  suggests that electron emission is likely to be a multiple photon process at visible wavelengths.

In addition to the electron emission energetics, we also consider the dissociation energetics, which are summarised in Table 2 for a range of possible channels involving the  $\text{MnO}_2^-$  product. Our calculations indicate that photodissociation is a lower energy pathway than photodetachment of  $\text{MnO}_4^-$  and suggest that photodissociation is likely to be a two-photon process for  $3.31$  eV  $< h\nu < 1.65$  eV ( $375$  nm  $< \lambda < 750$  nm). The lowest energy photodissociation channel is predicted to form  $\text{MnO}_2^- + \text{O}_2$  in their respective ground electronic states. The lowest energy  $\text{MnO}_3^- + \text{O}$  asymptote lies 0.4 eV higher in energy (3.7 eV). We do not include this channel in Table 2 because we have not observed  $\text{MnO}_3^-$  in our experiments, but we will return to this channel later. To the best of our knowledge, there are no reported experimental measurements of the dissociation asymptotes for comparison. Table 2 shows that the ground state of  $\text{MnO}_2^-$  is a quintet state ( $^5\text{B}_2$ ), with the triplet state ( $^3\text{B}_1$ ) lying 0.4 eV higher in energy. The lowest energy channels to form singlet states of either  $\text{MnO}_2^-$  or  $\text{O}_2$  are much higher in energy (note that the 1 eV gap between  $\text{O}_2$  singlet and

**Table 1** The relative adiabatic energetics, including zero point energies, of photodetachment and photodissociation of  $\text{MnO}_4^-$  and its anionic fragments

Process	Experimental energy (eV)	Theoretical energy (eV)
$\text{MnO}_4^- \rightarrow \text{MnO}_4 + \text{e}^-$	4.8 <sup>12</sup>	4.9
$\text{MnO}_4^- \rightarrow \text{MnO}_3^- + \text{O}$	—	3.7
$\text{MnO}_4^- \rightarrow \text{MnO}_2^- + \text{O}_2$	—	3.3
$\text{MnO}_2^- \rightarrow \text{MnO}_2 + \text{e}^-$	2.06 <sup>11</sup>	2.4
$\text{MnO}_3^- \rightarrow \text{MnO}_3 + \text{e}^-$	3.34 <sup>11</sup>	3.5

**Table 2** Relative adiabatic energetics, including zero point energies, of the different spin states of the dissociation asymptotes ( $D_0$ ) of  $\text{MnO}_4^-$

Dissociation asymptote	Relative energy (eV)
$\text{MnO}_2^- (^5\text{B}_2) + \text{O}_2 (^3\Sigma_g^-)$	3.3
$\text{MnO}_2^- (^3\text{B}_1) + \text{O}_2 (^3\Sigma_g^-)$	3.7
$\text{MnO}_2^- (^5\text{B}_2) + \text{O}_2 (^1\Delta_g)$	4.3
$\text{MnO}_2^- (^3\text{B}_1) + \text{O}_2 (^1\Delta_g)$	4.7
$\text{MnO}_2^- (^1\text{A}_1) + \text{O}_2 (^3\Sigma_g^-)$	4.9
$\text{MnO}_2^- (^1\text{A}_1) + \text{O}_2 (^1\Delta_g)$	5.9

triplet dissociation channels is consistent with the energy  $^3\Sigma_g^- - ^1\Delta_g$  difference of  $\text{O}_2$ ).<sup>27</sup>

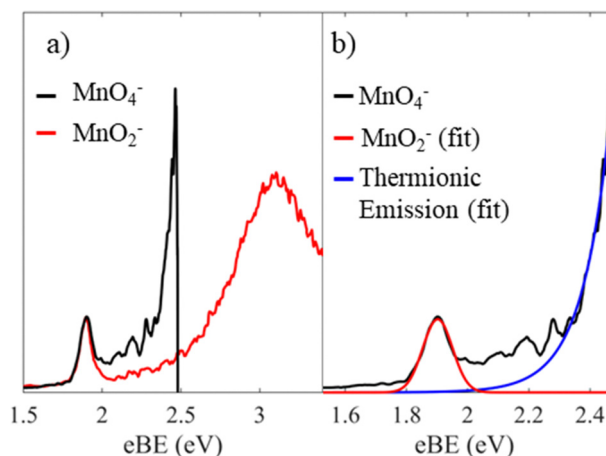
## Discussion

### (a) Assignment of triplet $\text{MnO}_2^-$ as photodissociation product

We first consider the assignment of the features in the photoelectron spectra in Fig. 2. Given the electron affinity of  $\text{MnO}_4^-$  is  $4.8 \pm 0.1$  eV, all the photoelectrons in Fig. 2 arise *via* multiple photon processes.<sup>12</sup> Light around  $h\nu \sim 2.5$  eV ( $\lambda \sim 500$  nm) is resonant with the bright  $^1\text{A}_1 \rightarrow ^1\text{T}_2$  transition, and given the rapid non-radiative decay observed in aqueous solution from the  $^1\text{T}_2$  state back to the ground state, thermionic emission is likely to occur *via* photon cycling.<sup>7,10,28,29</sup> In such a scenario, a second photon excites the hot-ground state *via* the same  $^1\text{A}_1 \rightarrow ^1\text{T}_2$  transition to leave  $\text{MnO}_4^-$  in the  $^1\text{A}_1$  ground state with sufficient internal energy to statistically dislodge the electron (*i.e.* internal energy is  $2 \times h\nu$ ). This thermionic emission feature was not seen in the previously reported photoelectron spectrum of  $\text{MnO}_4^-$ , because this spectrum was measured at a much higher photon energy ( $h\nu = 6.39$  eV,  $\lambda = 194$  nm) that is not resonant with the  $^1\text{A}_1 \rightarrow ^1\text{T}_2$  transition, and because of the low detection efficiency for low eKE electrons in their electron spectrometer (based on a magnetic bottle).<sup>12</sup>

The second narrow feature can be assigned to photodetachment from an excited state of  $\text{MnO}_2^-$ , as both the binding energy and the shape of the feature are consistent with the previously reported spectrum of  $\text{MnO}_2^-$ .<sup>11,30,31</sup> In Fig. 4(a), we overlay the photoelectron spectrum of Gustev *et al.* taken at  $h\nu = 3.49$  eV ( $\lambda = 355$  nm) with our spectrum of  $\text{MnO}_4^-$  taken at  $h\nu = 2.48$  eV ( $\lambda = 500$  nm), which clearly shows that the peak at eBE = 1.9 eV arises from  $\text{MnO}_2^-$ . However, the peak at eBE = 1.9 eV (Fig. 4(a)) is not the ground electronic state of  $\text{MnO}_2^-$ ; temperature dependent measurements by Gutsev *et al.* showed that this feature in fact arises from an excited state anion, which they assigned to the  $^3\text{B}_1$  state.<sup>11</sup> Hence, we conclude from the photoelectron spectrum in Fig. 2 that photodissociation leaves at least some of the  $\text{MnO}_2^-$  product in the  $^3\text{B}_1$  excited state.<sup>11</sup> It should also be noted that  $\text{MnO}_2^-$  may lose electrons *via* thermionic emission and contribute to the low eKE signal observed. None of the previously reported photoelectron spectra of  $\text{MnO}_2^-$  showed thermionic emission, although these were insensitive to low eKE electrons.<sup>11</sup>

Our measurements contain additional information in the form of PADs, which report on the nature of the molecular



**Fig. 4** Photoelectron spectrum of  $\text{MnO}_4^-$  recorded at  $h\nu = 2.48$  eV (black), where the energy axis is in electron binding energy,  $\text{eBE} = h\nu - \text{eKE}$ . In (a) this is shown with the previously reported spectrum of  $\text{MnO}_2^-$  taken at  $h\nu = 3.49$  eV (red) reprinted from *J. Chem. Phys.*, 2000, **113**, 1473, with the permission of AIP publishing.<sup>11</sup> In (b), a Gaussian fit to the spectral features associated with the  $\text{MnO}_2^-$  triplet state (red) is shown as well as an exponential fit to the thermionic emission (blue).

orbital that the photoelectron is ejected from. In this case, photodetachment from the excited state of  $\text{MnO}_2^-$  results in photoelectrons with  $\beta_2 \approx +1.0$ . Given that photodissociation and photodetachment of  $\text{MnO}_2^-$  is a multiple photon process of distinct sequential steps and occurs under irradiation with nanosecond light ( $\sim 5$  ns pulse duration), it is probable that the timescale of the process is much longer than rotational period of  $\text{MnO}_2^-$  and therefore that any potential alignment effects from the photodissociation process itself, can be ignored when interpreting the PADs. By Koopmans' theory, photodetachment results in ejection of an electron from the highest lying molecular orbital (HOMO), suggesting that the HOMO has s- or  $\sigma$ -orbital character (given the outgoing wave of p-character). We can use the simple d-block model of transition metal complexes to estimate the dominant contributions to the highest lying molecular orbitals, which are typically the non-bonding or antibonding metal-localised d-orbitals.<sup>32</sup> By considering the interaction of Mn with the two O atoms in the bent geometry, the  $d_{z^2}$  would be non-bonding, whereas the other d-orbitals would be higher in energy and antibonding in character. As a result of interactions between the O localised p-orbitals and the Mn localised  $d_{xz}$ ,  $d_{yz}$  and  $d_{x^2-y^2}$  orbitals, there would be three  $\pi^*$  orbitals, and a  $\sigma^*$  orbital from the interaction of  $d_{xy}$  with s- and p-orbitals orientated along the Mn–O bonds. These high-lying orbitals are then populated with the valence shell electrons of the metal, *i.e.*  $d^4$  for  $\text{Mn(III)}$ .<sup>32</sup> However for bent  $\text{ML}_2$  complexes, a metal localised p-orbital and a s–p hybrid are strictly non-bonding and higher in energy than the d-block, and as such these four electrons also populate the d-block.<sup>32</sup> The energy level diagram which arise from the d-block model for the triplet  $\text{MnO}_2^-$  is shown in Fig. 5.

Our calculations, which agree with previous work by Gutsev *et al.*, indicate that  $\text{MnO}_2^-$  preferentially adopts high spin electron configurations, *i.e.* the ground electronic state of



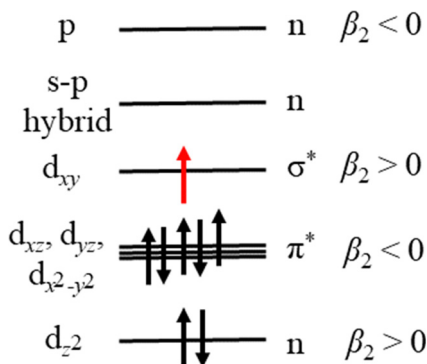


Fig. 5 An energy level diagram of the d-block of triplet  $\text{MnO}_2^-$ , including the expected  $\beta_2$  values if the electron is lost. The molecular orbitals are populated with the  $d^4$  electrons of  $\text{Mn(III)}$  and 4 electrons from the high lying non-bonding orbitals of a bent  $\text{ML}_2$  complex, whilst the electron, which is lost in photodetachment, via Koopmans' correlations, is highlighted red. Expected anisotropy parameters are indicated.

$\text{MnO}_2^-$  is  $^5\text{B}_2$ .<sup>11</sup> Effectively this suggests that the pairing energy to doubly occupy d-block orbitals is higher than the ligand field splitting in this case. Given the temperature dependent measurements of Gutsev *et al.*, we know that photodetachment of  $\text{MnO}_2^-$  arises from an excited state *i.e.* a singlet or triplet.<sup>11</sup> We have populated the d-block in Fig. 5 with the electron configuration of the  $^3\text{B}_1$  state of  $\text{MnO}_2^-$ , which calculations have consistently determined to be the first excited state (Table 2 and Gutsev *et al.*<sup>11</sup>). If photodetachment removed the red electron from the HOMO in Fig. 5, which is a  $\sigma^*$  molecular orbital, it would be expected to result in a PAD characterised by  $\beta_2 > 0$ , as observed experimentally and consistent with the assignment that  $\text{MnO}_2^-$  is produced in the  $^3\text{B}_1$  electronically excited state. If instead detachment took place from the lowest energy singlet electronic state, then we would expect the PAD to be characterised by  $\beta_2 < 0$  on account of its  $\pi$ -character, which is not observed.

It is more challenging to determine if ground state  $\text{MnO}_2^-$  is also produced. In the photoelectron spectrum of  $\text{MnO}_2^-$  by Gutsev *et al.* (Fig. 4(a)), photoelectrons with  $\text{eBE} > 2.1$  eV were assigned to photodetachment from the ground  $^5\text{B}_2$  state of  $\text{MnO}_2^-$ .<sup>11</sup> If ground state  $\text{MnO}_2^-$  was produced in our experiments, then the associated spectral features would overlap with the thermionic emission. In order to disentangle these potential contributions, we fit our spectrum with a Gaussian and exponential decay to account for the direct and thermionic detachment channels, respectively, as shown in Fig. 4(b). While the fit accounts for most signal, some signal is unaccounted for. However, the spectrum of the unaccounted for electrons does not match the ground state spectrum of  $\text{MnO}_2^-$ . Furthermore, the PAD in this region is isotropic, which would be very unusual for a transition metal complex.<sup>33,34</sup> From Fig. 5 we may expect that the  $^5\text{B}_2$  state would lose an electron from the  $s-p$  hybrid molecular orbital, and would be expected to exhibit an anisotropic PAD. We attempted to verify this by recording the photoelectron spectrum of  $\text{MnO}_2^-$  directly but were unsuccessful as we could not produce  $\text{MnO}_2^-$  *via* electrospray ionization.

Therefore, it is likely that the mismatch between the fits and the recorded data is either due to noise, which may be present because of the very low data acquisition rates for our photoelectron spectra, or due to additional thermionic emission leading to an emission profile that cannot be accounted for by a single exponential. Overall, we see no direct evidence for the formation of  $\text{MnO}_2^-$  in its electronic ground state. However, given that the vertical detachment energy of ground state  $\text{MnO}_2^-$  is near  $\text{eBE} \sim 3.3$  eV (see Fig. 4(a)), and the low signal levels in our experiments, it is not possible to conclusively rule it out. Ideally, we would like to have performed two-colour experiments, with a UV probe pulse, in order to search for ground state  $\text{MnO}_2^-$ , but the signal levels were too low for these measurements to be successful.

### (b) Assignment of electron action spectrum

Our electron action spectrum of  $\text{MnO}_4^-$  and the previously reported  $\text{MnO}_2^-$  action spectrum of  $\text{MnO}_4^-$  by Houmoller *et al.* are plotted together in Fig. 6.<sup>7</sup> Both spectra show an identical vibrational progression, associated with the Mn–O stretch vibrational mode of  $\text{MnO}_4^-$  in the  $1^1\text{T}_2$  state and observed in the absorption spectrum of aqueous  $\text{MnO}_4^-$ .<sup>1</sup> The similarity indicates that both photodissociation to produce  $\text{MnO}_2^-$  and thermionic emission involve photoexcitation of the  $1^1\text{T}_2$  state of  $\text{MnO}_4^-$ . However, there are some clear differences too. Specifically, the first three vibrational peaks appear to be missing in the electron action spectrum. These differences arise as there

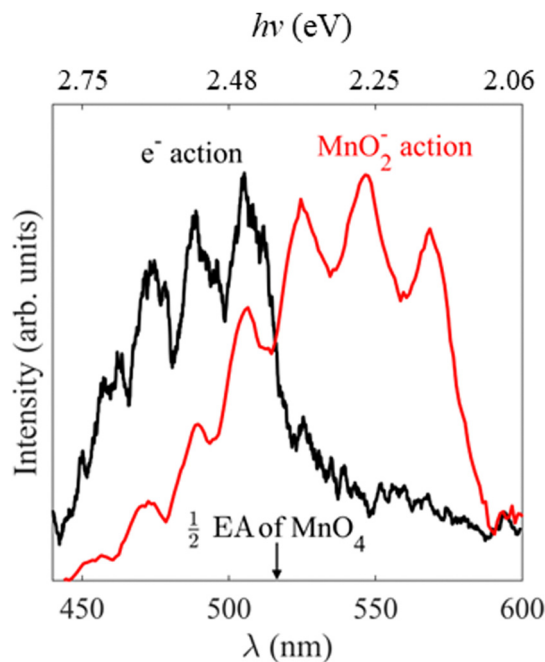


Fig. 6 Electron action spectrum of  $\text{MnO}_4^-$  recorded between  $h\nu = 2.82$  and  $2.07$  eV ( $\lambda = 440$  and  $600$  nm) in black, with the previously reported  $\text{MnO}_2^-$  action spectrum of  $\text{MnO}_4^-$  in red, reprinted from *Chem. Phys. Chem.*, 2013, **14**, 1133 with the permission of John Wiley and Sons.<sup>7</sup> The wavelength at which two-photon photodetachment of  $\text{MnO}_4^-$  becomes energetically accessible *i.e.*, half the previously reported EA of  $\text{MnO}_4^-$ , is indicated by the black arrow.<sup>12</sup>



are different photoprocesses contributing to each of the action spectra. Photodissociation to produce  $\text{MnO}_2^-$  is the only contributor to the ion action spectrum, whereas both photodetachment from  $\text{MnO}_2^-$  as well as thermionic emission (from either  $\text{MnO}_2^-$  or  $\text{MnO}_4^-$ ) contribute to the electron action of  $\text{MnO}_4^-$  (see Fig. 2).

Our calculations indicate that  $\text{MnO}_2^- (^3\text{B}_1) + \text{O}_2 (^3\Sigma_g^-)$  lies 3.7 eV above  $\text{MnO}_4^-$ , requiring two visible photons with  $h\nu > 1.85$  eV ( $\lambda < 670$  nm) to photodissociate to our observed products. Houmoller *et al.* confirm that the photodissociation is a two-photon process at these wavelengths by performing fluence dependent measurements and suggest that the dissociation asymptote is  $< 4$  eV, in accordance with our calculations.<sup>7</sup> In our experiments a further photon is required to photodetach an electron from  $\text{MnO}_2^-$ , which we subsequently detect in our electron action spectrum, making this channel a three-photon process overall. From our photoelectron spectra in Fig. 2, the triplet state of  $\text{MnO}_2^-$  has an eBE  $\sim 1.9$  eV (two-photons at  $\lambda = 650$  nm), meaning that the three-photon process is energetically accessible at all the wavelengths studied here, as long as a dissociation takes place on a timescale sufficiently short that the third photon can be observed during the laser pulse ( $\sim 5$  ns).

In addition, thermionic emission contributes to the  $\text{MnO}_4^-$  electron action spectrum, whether from  $\text{MnO}_4^-$  directly or the photodissociation product of  $\text{MnO}_2^-$ . The electron affinity of  $\text{MnO}_4^-$  is 4.8 eV, requiring a three-photon process at the longest wavelengths studied, but a two-photon electron loss process would become energetically accessible at  $h\nu \sim 2.4$  eV ( $\lambda \sim 516$  nm). If photodissociation is the dominant product channel following excitation of  $\text{MnO}_4^-$  at visible wavelengths, then the thermionic emission may also arise from  $\text{MnO}_2^-$  via a three-photon process.

A large increase in signal is seen in the electron action spectrum near  $h\nu \sim 2.4$  eV ( $\lambda \sim 520$  nm), but the origin of this increase is not clear cut. From Fig. 2 it is seen that this increase coincides somewhat with a  $\sim 75\%$  increase in laser power, which would be expected to rapidly increase the likelihood of multiple photon processes (the rate of an  $n$ -photon process is typically  $\propto P^n$ , where  $P$  is the laser fluence). Alternatively, electron loss from  $\text{MnO}_4^-$  switches from a two-to-three photon transition for  $h\nu > 2.4$  eV ( $\lambda < 516$  nm), which is indicated in Fig. 6, and also coincides with the increase in overall electron signal. In order to test these different hypotheses, we performed additional measurements which were inconclusive. Fluence dependent measurements (Fig. S1, ESI†) at  $h\nu = 2.46$  eV ( $\lambda = 505$  nm) showed no difference in the ratio of direct detachment to thermionic emission with laser power, indicating that dynamics of both photodissociation and photodetachment are limited by the same step, likely to be the initial photoexcitation from the  $^1\text{A}_1 \rightarrow ^1\text{T}_2$  transition. Furthermore, the photoelectron spectrum of  $\text{MnO}_4^-$  at  $h\nu = 2.34$  eV ( $\lambda = 530$  nm) (Fig. S1, ESI†), where only three-photon processes are energetically accessible, unexpectedly showed a higher proportion of thermionic emission compared to direct detachment from  $\text{MnO}_2^-$ , than spectra recorded at shorter

wavelengths. Ultimately it is probable that a combination of increased laser power and electron loss from  $\text{MnO}_4^-$  switching from a three- to a two-photon process, contribute to the shape of the electron action spectrum. However, it is challenging to unravel comprehensively, as these factors are seen against a backdrop of competing multiple photon photoprocesses, with different cross-sections for each step at each wavelength.

### (c) Photodissociation mechanism

The production of  $\text{MnO}_2^-$  in an excited state must result directly from the photodissociation process. This preference may be due to the requirement to conserve spin during bond breaking. Given that the ion action spectrum of Houmoller *et al.* indicated that  $\text{MnO}_2^-$  was produced *via* a two-photon process at  $h\nu = 4.27$  eV ( $2 \times \lambda = 580$  nm), only the lowest two energy asymptotes from Table 2 are energetically accessible, with the third being within the error of the calculations.

The requirement to conserve spin is applied to the final electronic state of  $\text{MnO}_4^-$  before dissociation, and  $\text{MnO}_4^-$  has singlet, triplet, quintet and septet electronic states.<sup>8</sup> Removal of two electrons from the bonding orbitals of  $\text{MnO}_4^-$  with spin conservation, to account for the breaking of two Mn–O bonds during photodissociation, allows us to predict the spin of the products. This process is illustrated in Fig. 7 for the energetically accessible dissociation asymptotes and demonstrates that the spin of the parent anion dictates the spin state of the products (in the absence of spin–orbit coupling). Therefore, by conservation of spin angular momentum, the observed  $\text{MnO}_2^- (^3\text{B}_1)$  can only arise from dissociation of a singlet  $\text{MnO}_4^-$  state, whereas ground state  $\text{MnO}_2^- (^5\text{B}_2)$  may be produced *via* dissociation of a triplet state of  $\text{MnO}_4^-$ .

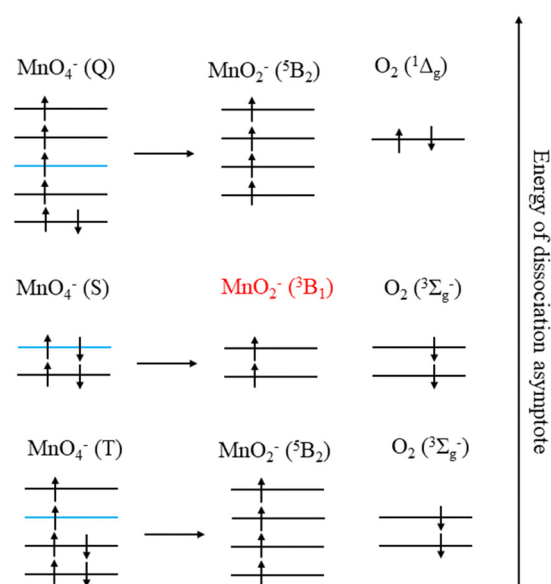
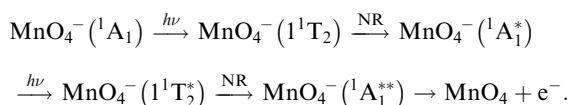


Fig. 7 The singlet (S), triplet (T) and quintet (Q) spin states of  $\text{MnO}_4^-$  which can produce the three energetically accessible dissociation asymptotes of  $\text{MnO}_2^- + \text{O}_2$ . The experimentally observed spin state of  $\text{MnO}_2^-$  is highlighted in red, with the  $\text{MnO}_4^-$  singlet HOMO being highlighted in blue for all the spin states.

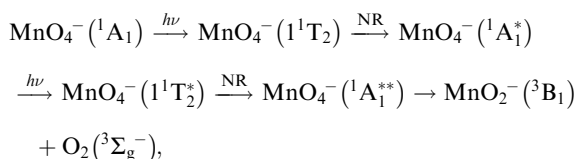


From spin conservation only, the singlet state of  $\text{MnO}_4^-$  upon which dissociation occurs may be either the ground state ( $^1\text{A}_1$ ), the excited state accessed at visible wavelengths ( $^1\text{T}_2$ ) or the  $^1\text{T}_1$  transient intermediate that may be populated in the decay mechanism of the  $^1\text{T}_2$  state (see Fig. 1). Evidence for a ground state dissociation of  $\text{MnO}_4^-$  is found in the presence of thermionic emission in the photoelectron spectra, the structure of our electron action spectrum, and the previously reported ion action spectrum of  $\text{MnO}_4^-$ .<sup>7</sup> Therefore, the overall mechanism of non-radiative (NR) decay from  $^1\text{T}_2$  to  $^1\text{A}_1$  is likely to be the dominant decay pathway, as observed in aqueous solution to take place on a 16 ps timescale.<sup>4</sup> The vibrational structure in the both the electron and the ion action spectrum of  $\text{MnO}_4^-$  also indicates that both photodetachment and photodissociation are initiated by the same  $^1\text{A}_1 \rightarrow ^1\text{T}_2$  transition, and again provides evidence for a ground state mechanism.<sup>7</sup> A cartoon of the two-photon process which results in a vibrationally hot ground state  $\text{MnO}_4^-$ , and acts as the precursor to electron loss or dissociation, is shown in Fig. 8.

The overall two-photon process for thermionic emission can be described as follows, where the \* represents a photon's worth of internal excitation (*i.e.* heat):

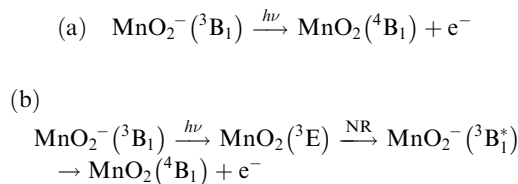


An analogous two-photon ground state dissociation mechanism for the production of  $\text{MnO}_2^-$  may also be proposed:



where the  $\text{MnO}_2^-$  product loses electrons by absorption of a further photon (three in total) either *via* (a) direct detachment

or (b) thermionic emission (mediated by an unknown excited state,  $^3\text{E}$ ):



It is clear from our measurements that  $\text{MnO}_2^- (^3\text{B}_1)$  is a long-lived electronically excited state, with a lifetime on the order of ns, which does not readily return to the ground state. This must be the case as we record photodetachment from the excited state, indicating that the absorption of a third photon occurs on a faster, or at least similar, timescale to excitation (laser pulse is  $\sim 5$  ns in duration). Furthermore Gutsev *et al.* observed some  $\text{MnO}_2^- (^3\text{B}_1)$  directly in their anion beam, which suggests a lifetime on the order of their time of flight, which is many  $\mu\text{s}$ .<sup>11</sup>

In aqueous solution, excitation to the  $^1\text{T}_2$  band also shows the formation of  $\text{MnO}_2^-$ , although it is not known in what electronic state. It is also produced in a very low yield. These differences are likely associated with the fact that the hot ground state formed after non-radiative decay can dissipate its internal energy before absorbing a second photon, inhibiting dissociation.

Finally, we also consider the  $\text{MnO}_3^- + \text{O}$  channel. Previous gas phase ion action spectroscopy has reported the production of  $\text{MnO}_3^-$  from irradiation of  $\text{MnO}_4^-$  in the visible region. We were unable to observe any  $\text{MnO}_3^-$  produced *via* photodissociation of  $\text{MnO}_4^-$  in our experiments, as the electron affinity ( $\sim 3.1$  eV) is larger than the photon energies used ( $2.0 < h\nu < 2.75$  eV). However, it was observed as an almost equal yield channel as the  $\text{MnO}_2^- + \text{O}_2$  channel by Houmoller *et al.* and this result is consistent with our data. The onset for the  $\text{MnO}_3^- + \text{O}$  channel is 0.4 eV above that of the lowest  $\text{MnO}_2^- (^5\text{B}_2) + \text{O}_2 (^3\Sigma_g^-)$  channel. Hence, statistically, the former should be much less likely. However, as the  $\text{MnO}_2^- (^5\text{B}_2) + \text{O}_2 (^3\Sigma_g^-)$  channel is closed because of the required spin-flip, the relevant channel is the  $\text{MnO}_2^- (^3\text{B}_1) + \text{O}_2 (^3\Sigma_g^-)$  dissociation channel, which is 0.4 eV higher in energy and essentially degenerate with the  $\text{MnO}_3^- + \text{O}$  channel (according to our calculations). It is interesting to note that the lowest energy dissociation asymptote is predicted to be  $\text{MnO}_3^- (^1\text{A}_1') + \text{O} (^3\text{P}_2)$ , which would only be spin-allowed from a triplet state of  $\text{MnO}_4^-$  (Fig. 7). However, calculations have indicated that there is a low-lying triplet state of  $\text{MnO}_3^-$ , which using our theoretical methodology lies 0.1 eV above the ground state, and could be produced with conservation of spin from a singlet state of  $\text{MnO}_4^-$  *i.e.*, *via* a ground state dissociation.<sup>11</sup> Ideally two-colour pump-probe measurements to probe the production of  $\text{MnO}_3^-$  would be undertaken. However, the attempted time-resolved measurements to study the production of  $\text{MnO}_2^-$  were unsuccessful, due to the very low signal levels observed as a result of the two-photons required for photodissociation. Note also that

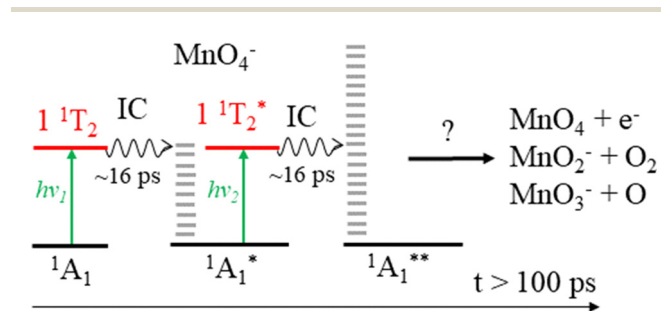


Fig. 8 Sequential absorption of two photons near  $h\nu = 2.48$  eV ( $\lambda = 500$  nm) by  $\text{MnO}_4^-$  produces a highly internally excited  $\text{MnO}_4^-$  in its electronic ground state (where \* indicates a photon's worth of internal excitation), which subsequently can lose either an electron, an O atom or  $\text{O}_2$  *via* a statistical process. The IC process involves several nonradiative transitions as shown in Fig. 1.



$\text{MnO}_3^-$  was not observed as a product in aqueous solution, which might indicate that dissociation in solution is preferred to proceed to the ground state product of  $\text{MnO}_2^-$ , where perhaps a spin-flip is mediated by collisions with solvent molecules.

## Conclusions

It has been known for many decades that aqueous permanganate photodissociates to produce  $\text{MnO}_2^-$  following excitation with green light, but the mechanism has remained the subject of active research.<sup>3</sup> In this work we demonstrate that the gas-phase photodissociation of  $\text{MnO}_4^-$  in the visible region is a multiple photon ground state process, which produces  $\text{MnO}_2^-$  in an electronically excited triplet state. Our proposed mechanism of ground state dissociation is consistent with recent femtosecond transient absorption spectroscopy in the solution phase which unravelled the rapid non-radiative decay dynamics of  $\text{MnO}_4^-$ .<sup>4</sup> Furthermore, it may explain why  $\text{MnO}_3^-$  has been observed as a photodissociation product of  $\text{MnO}_4^-$  in the gas phase, but not the solution phase.<sup>7</sup> Our conclusions are reached *via* photoelectron imaging, electron action spectroscopy and electronic structure calculations, where competition is observed between a statistical photodissociation and ground state electron loss from  $\text{MnO}_4^-$ , *via* thermionic emission.

## Data availability statement

The data that support the findings of this study are available in Zenodo at <https://zenodo.org/records/10204163>.

## Conflicts of interest

There are no conflicts of interest to declare.

## Acknowledgements

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## References

- 1 S. L. Holt and C. J. Ballhausen, Low temperature absorption spectra of  $\text{KMnO}_4$  in  $\text{KClO}_4$ , *Theor. Chim. Acta*, 1967, **7**(4), 313–320.
- 2 J. H. Mathews and L. H. Dewey, A quantitative study of some photochemical effects produced by ultra-violet light, *J. Phys. Chem.*, 1913, **17**(3), 211–218.
- 3 G. Zimmerman, Photochemical Decomposition of Aqueous Permanganate Ion, *J. Chem. Phys.*, 1955, **23**(5), 825–832.
- 4 O. S. Haggag, P. Malakar, P. Pokhilko, J. F. Stanton, A. I. Krylov and S. Ruhman, The elusive dynamics of aqueous permanganate photochemistry, *Phys. Chem. Chem. Phys.*, 2020, **22**(18), 10043–10055.
- 5 A. W. Adamson, W. L. Waltz, E. Zinato, D. W. Watts, P. D. Fleischauer and R. D. Lindholm, Photochemistry of transition-metal coordination compounds, *Chem. Rev.*, 1968, **68**(5), 541–585.
- 6 D. G. Lee, C. R. Moylan, T. Hayashi and J. I. Brauman, Photochemistry of aqueous permanganate ion, *J. Am. Chem. Soc.*, 1987, **109**(10), 3003–3010.
- 7 J. Houmøller, S. H. Kaufman, K. Støchkel, L. C. Tribedi, S. Brøndsted Nielsen and J. M. Weber, On the Photoabsorption by Permanganate Ions in Vacuo and the Role of a Single Water Molecule. New Experimental Benchmarks for Electronic Structure Theory, *ChemPhysChem*, 2013, **14**(6), 1133–1137.
- 8 G. L. Gutsev, B. K. Rao and P. Jena, Photodecomposition of  $\text{MnO}_4^-$ : A Theoretical Study, *J. Phys. Chem. A*, 1999, **103**(50), 10819–10824.
- 9 H. Nakai, Y. Ohmori and H. Nakatsuji, Theoretical Study on the Photochemical Decomposition Reaction of Permanganate Ion,  $\text{MnO}_4^-$ , *J. Phys. Chem.*, 1995, **99**(21), 8550–8555.
- 10 L. Jose, M. Seth and T. Ziegler, Molecular and Vibrational Structure of Tetroxo  $d^0$  Metal Complexes in their Excited States. A Study Based on Time-Dependent Density Functional Calculations and Franck–Condon Theory, *J. Phys. Chem. A*, 2012, **116**(7), 1864–1876.
- 11 G. L. Gutsev, B. K. Rao, P. Jena, X. Li and L.-S. Wang, Experimental and theoretical study of the photoelectron spectra of  $\text{MnO}_x^-$  ( $x = 1-3$ ) clusters, *J. Chem. Phys.*, 2000, **113**(4), 1473–1483.
- 12 G. L. Gutsev, B. K. Rao, P. Jena, X.-B. Wang and L.-S. Wang, Origin of the unusual stability of  $\text{MnO}_4^-$ , *Chem. Phys. Lett.*, 1999, **312**(5), 598–605.
- 13 J. Lecointre, G. M. Roberts, D. A. Horke and J. R. R. Verlet, Ultrafast Relaxation Dynamics Observed Through Time-Resolved Photoelectron Angular Distributions, *J. Phys. Chem. A*, 2010, **114**(42), 11216–11224.
- 14 L. H. Stanley, C. S. Anstöter and J. R. R. Verlet, Resonances of the anthracenyl anion probed by frequency-resolved photoelectron imaging of collision-induced dissociated anthracene carboxylic acid, *Chem. Sci.*, 2017, **8**(4), 3054–3061.
- 15 W. C. Wiley and I. H. McLaren, Time-of-Flight Mass Spectrometer with Improved Resolution, *Rev. Sci. Instrum.*, 1955, **26**(12), 1150–1157.
- 16 G. M. Roberts, J. L. Nixon, J. Lecointre, E. Wrede and J. R. R. Verlet, Toward real-time charged-particle image reconstruction using polar onion-peeling, *Rev. Sci. Instrum.*, 2009, **80**(5), 053104.
- 17 A. Sanov, Laboratory-Frame Photoelectron Angular Distributions in Anion Photodetachment: Insight into Electronic Structure and Intermolecular Interactions, *Annu. Rev. Phys. Chem.*, 2014, **65**(1), 341–363.
- 18 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov,



- J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery Jr., J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman and D. J. Fox, *Gaussian 16 Rev. C.01*, Wallingford, CT, 2016.
- 19 A. D. Becke, Density-functional thermochemistry. III. The role of exact exchange, *J. Chem. Phys.*, 1993, **98**(7), 5648–5652.
  - 20 R. A. Kendall, T. H. Dunning Jr. and R. J. Harrison, Electron affinities of the first-row atoms revisited. Systematic basis sets and wave functions, *J. Chem. Phys.*, 1992, **96**(9), 6796–6806.
  - 21 J. A. Gibbard, J. Reppel and J. R. R. Verlet, Photoelectron imaging of  $\text{MnO}_3^-$  to probe its nuclear and electronic structure, *J. Chem. Phys.*, 2023, **159**(14), 146101.
  - 22 B. Baguenard, J. C. Pinaré, C. Bordas and M. Broyer, Photoelectron imaging spectroscopy of small tungsten clusters: Direct observation of thermionic emission, *Phys. Rev. A: At., Mol., Opt. Phys.*, 2001, **63**(2), 023204.
  - 23 B. Baguenard, J. C. Pinaré, F. Lépine, C. Bordas and M. Broyer, Thermionic emission in small carbon cluster anions, *Chem. Phys. Lett.*, 2002, **352**(3), 147–153.
  - 24 K. Hansen, K. Hoffmann and E. E. B. Campbell, Thermal electron emission from the hot electronic subsystem of vibrationally cold  $\text{C}_{60}$ , *J. Chem. Phys.*, 2003, **119**(5), 2513–2522.
  - 25 M. Kjellberg, O. Johansson, F. Jonsson, A. V. Bulgakov, C. Bordas, E. E. B. Campbell and K. Hansen, Momentum-map-imaging photoelectron spectroscopy of fullerenes with femtosecond laser pulses, *Phys. Rev. A: At., Mol., Opt. Phys.*, 2010, **81**(2), 023202.
  - 26 K. P. Huber and G. Herzberg, Constants of diatomic molecules, in *Molecular Spectra and Molecular Structure: IV. Constants of Diatomic Molecules*, ed. K. P. Huber and G. Herzberg, Springer US, Boston, MA, 1979, pp. 8–689.
  - 27 C. Schweitzer and R. Schmidt, Physical Mechanisms of Generation and Deactivation of Singlet Oxygen, *Chem. Rev.*, 2003, **103**(5), 1685–1758.
  - 28 J. N. Bull, C. W. West and J. R. R. Verlet, On the formation of anions: frequency-, angle-, and time-resolved photoelectron imaging of the menadione radical anion, *Chem. Sci.*, 2015, **6**(2), 1578–1589.
  - 29 J. A. Gibbard, C. J. Clarke and J. R. R. Verlet, Photoelectron spectroscopy of the protoporphyrin IX dianion, *Phys. Chem. Chem. Phys.*, 2021, **23**(34), 18425–18431.
  - 30 K. M. Ervin, I. Anusiewicz, P. Skurski, J. Simons and W. C. Lineberger, The Only Stable State of  $\text{O}_2^-$  Is the  $X^2\Pi_g$  Ground State and It (Still!) Has an Adiabatic Electron Detachment Energy of 0.45 eV, *J. Phys. Chem. A*, 2003, **107**(41), 8521–8529.
  - 31 S. J. Cavanagh, S. T. Gibson, M. N. Gale, C. J. Dedman, E. H. Roberts and B. R. Lewis, High-resolution velocity-map-imaging photoelectron spectroscopy of the  $\text{O}^-$  photodetachment fine-structure transitions, *Phys. Rev. A: At., Mol., Opt. Phys.*, 2007, **76**(5), 052708.
  - 32 Y. Jean, *Molecular orbitals of transition metal complexes*, OUP, Oxford, UK, 2005.
  - 33 J. A. Gibbard and J. R. R. Verlet, Photoelectron Imaging Study of the Diplatinum Iodide Dianions  $[\text{Pt}_2\text{I}_6]^{2-}$  and  $[\text{Pt}_2\text{I}_8]^{2-}$ , *J. Phys. Chem. A*, 2022, **126**(22), 3495–3501.
  - 34 J. A. Gibbard and J. R. R. Verlet, Photoelectron imaging of  $\text{PtI}_2^-$  and its  $\text{PtI}^-$  photodissociation product, *J. Chem. Phys.*, 2022, **156**(13), 134303.

