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# The electronic structure of transition metal oxides for oxygen evolution reaction

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Electrolysis of water to produce hydrogen and oxygen is a promising pathway for the storage of renewable energy in form of chemical fuels. The efficiency of the overall process is usually limited by the sluggish kinetics of the oxygen evolution reaction (OER) due to a complex four-electron/proton transfer mechanism. Therefore, the most crucial step for water electrolysis to become a widespread industrial process is to develop efficient electrocatalysts capable of driving the OER at a low overpotential. In this article, we present a review of the recent understanding of low-cost transition metal oxide-based catalysts for electrochemical OER. Our main focus is on the fundamental understanding of the electronic structure of these materials and its relationship with the catalytic activity and the reaction mechanism. We detail the electronic structure descriptors of the OER performance that have provided guidelines for the development of advanced electrocatalysts. Strategies to enhance the activity based on such electronic structure guidelines, including doping, strain and defect engineering are also discussed. Finally, we summarize the existing challenges and present perspectives for the future development of the field.

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## 1. Introduction

Since the 19th century, non-renewable energy resources such as coal, oil, and natural gas have been utilized as the main energy

supply to sustain human activities.<sup>1</sup> However, the excessive consumption of fossil fuels caused by the rapidly increasing world population and expanding industrialization as well as concomitant environmental problems have motivated our society to explore a green and sustainable energy system for long-term development.<sup>2–5</sup> In this regard, countries around the world have made gigantic efforts to raise the share of renewable and carbon-free sources of energy, such as solar, wind, and tidal energy.<sup>6</sup> Unfortunately, these energy sources have a general intermittent availability and heavily depend on the season and weather, which result in a low energy delivery efficiency and

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restrict their widespread application. Thus, the storage of renewable energy in the form of fuels and useful chemicals is a highly valuable approach to introduce such sources of renewable energy in the mainstream of our current energy economy. In this regard, hydrogen has been considered as a promising energy carrier because it has zero-carbon combustion products and the highest gravimetric energy density.<sup>7,8</sup> Hydrogen not only can be used as energy carrier itself but it is also an important industrial feedstock for methanol production and ammonia synthesis among other processes. Since technologies for the conversion of many sources of renewable energy (e.g., solar irradiance and wind) into electrical power are well-developed these days, the electrochemistry provides means of storing the surplus of solar- or wind-driven electricity into fuels and valuable chemicals with a low carbon

footprint.<sup>9</sup> In a joint effort, scientists, policymakers and politicians around the world have carried extensive research and investment on projects for the large-scale wind power or photovoltaic hydrogen production, which combining the utilization of renewable energy with the hydrogen generation through water electrolysis.<sup>10</sup> Currently, the global hydrogen production reaches more than 110 Mt per year.<sup>11</sup> However, the H<sub>2</sub> production through water electrolysis only takes 4% of whole hydrogen supply,<sup>12</sup> because the high cost to drive the electrolytic water splitting (>US\$4 per kg of H<sub>2</sub>) compared with the hydrogen production cost through the reforming of fossil fuels (US\$1.3–1.5 per kg of H<sub>2</sub>).<sup>11</sup> The cost of industrial water electrolysis mainly arises from the large electric energy input required, which is closely related with the composition of applied catalysts and their catalytic performance.<sup>13</sup> Thus, developing cheap catalysts with high activities is significant to reduce the production cost with better economic benefits (<US\$2 per kg of H<sub>2</sub>) and broaden the commercial and industrial application of electrochemical water splitting.

Since Troostwijk and Diemann firstly observed the phenomenon of water electrolysis in 1789,<sup>14</sup> it has been extensively investigated by researchers over the past years. It is now well-established that water electrolysis involves two fundamental half reactions: hydrogen evolution reaction (HER) at the cathode, and the oxygen evolution reaction (OER) at the anode. Driving the electrochemical water splitting thermodynamically requires to overcome the energy barrier of 237.2 kJ mol<sup>-1</sup> at standard conditions, which corresponds to an applied voltage of 1.23 V in an electrochemical cell.<sup>15,16</sup> However, in practice, the actual applied cell voltage is much larger than the theoretical value due to the overpotentials required to drive both the HER and the OER.<sup>17</sup> As a result, even in well-developed and optimised water electrolysis systems, such as proton-exchange membrane water electrolyzers for electrolysis in acidic condition (Fig. 1), a voltage of 1.7–2.2 V is required to deliver a current density of >1.6 A cm<sup>-2</sup>.<sup>18</sup> The situation is similar in anion-



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Fig. 1 Schematics of the electrochemical water splitting in anion-exchange membrane electrolyser and proton-exchange membrane electrolyser.

exchange membrane electrolyzers for electrolysis in alkaline media (Fig. 1) in which a voltage of 1.7–2.4 V is usually required to deliver a current density of  $<1.0 \text{ A cm}^{-2}$ .<sup>19</sup> Most of the overpotential loss comes from the OER that involves a multi-step transfer process of electrons and protons with a complex reaction mechanism, resulting in sluggish kinetics and larger activation energy barriers than those associated to the HER process.<sup>20</sup> In this scenario, developing more active catalysts toward the OER has become the most important factor to improve the overall efficiency of water electrolysis for industrial applications.

After decades of efforts, the worldwide research for OER catalysts have made significant progress in promoting the catalytic activities. Numerous novel materials have emerged as advanced OER catalysts such as metal chalcogenide,<sup>21–23</sup> oxides,<sup>24,25</sup> nitrides<sup>26,27</sup> and molecular/polymeric systems.<sup>28,29</sup> Ru- and Ir-based oxides are OER catalysts of high industrial relevance because their high activity in acidic media and implementation in proton-exchange membrane electrolyzers.<sup>30,31</sup> However, both RuO<sub>2</sub> and IrO<sub>2</sub> undergo further oxidation and subsequent dissolution in acidic electrolytes under high applied potentials,<sup>32,33</sup> leading to unstable performance during water electrolysis. Furthermore, the high price of Ru and Ir also inhibit their large-scale application. On the other hand, one of the major advantages of alkaline water electrolysis over the acidic electrolysis is the replacement of conventional noble metal electrocatalysts with low-cost transition metal (TM) catalysts. Additionally, an alkaline media provides higher stability for the OER metal oxide catalysts.<sup>19,34</sup> Nevertheless, the anion-exchange membrane electrolysis is still a developing technology, and the current density and voltage efficiency achieved are lower than those in the electrolysis in acidic media. Therefore, considering the cost and activity for the practical utilization, extensive research efforts must be focused on investigating earth-abundant transition metal oxide (TMO) electrocatalysts including perovskite family,<sup>35,36</sup> spinel family,<sup>37–39</sup> layer-structured hydroxides<sup>40–42</sup> and other types of

oxides.<sup>43,44</sup> The OER performance of reported outstanding TMOs have exhibited remarkable activities and relatively high electrochemical stability, especially in alkaline electrolytes.<sup>45–47</sup>

In order to develop a competitive low-cost anion-exchange membrane electrolysis technology for commercially viable hydrogen production, further research and improvements are required, specifically regarding the reaction mechanisms on the surface of the catalysts and the development of guidelines for the preparation of advanced OER electrodes. Earlier in 1955, Rüttschi and Delahay reported an approximate linear relationship between the rate of the OER on metal oxides and the M–OH binding strength.<sup>48</sup> They made the pioneering attempt to link the OER activity with a descriptor. Furthermore, in the 1970s, Matsumoto *et al.*<sup>49</sup> firstly studied the correlation between the OER catalytic activity of perovskite oxides and their band structure. They proposed that the perovskite-type oxide formed  $\sigma^*$  and  $\pi^*$  band, and the rate of electron transfer would increase with the rise of the  $\sigma^*$  bandwidth. Later, Bockris and Otagawa systematically investigated the OER activities on perovskite oxides and they connected the OER activity with the number of electrons occupied in the antibonding orbitals.<sup>50</sup> Based on the band structure theory and molecular orbital theory, a few typical descriptors have been recently proposed to link the activity with electronic structure and thus theoretically predict the catalytic activity of novel catalysts. The most notable electronic descriptors of the OER activity include the number of electrons in the  $e_g$  orbitals of metal cations,<sup>51,52</sup> the degree of TM 3d–O 2p hybridization,<sup>53</sup> the energy position of the O 2p band center,<sup>54</sup> charge-transfer energy,<sup>35</sup> bulk thermochemistry<sup>55</sup> among others. For instance, Shao-Horn and co-workers systematically studied the relation between OER activities and the number of  $e_g$  electrons for perovskite oxides, conducting a conclusion that the optimal catalytic performance could be achieved when the number of electrons filled in  $e_g$  orbital was 1.2.<sup>51</sup> Later, Xu *et al.*<sup>56</sup> reported that  $e_g$  orbital occupancy correlates well with the OER activity of spinel type oxides. Zhou *et al.*<sup>57</sup> and Xie *et al.*<sup>58</sup> report that the activity can be optimized by adjusting the

spin state of Co 3d orbital electron in the LaCoO<sub>3</sub> system. Moreover, Grimaud *et al.*<sup>59</sup> also use the position of O 2p-band centre relative to the Fermi level as a descriptor of the OER

catalytic activities. The investigations on the correlation of the OER catalytic activity with electronic structure have provided meaningful results. A comprehensive understanding of these

Table 1 The performance of various TM oxide electrocatalysts for the OER

Electrocatalysts	Electrolyte	Catalyst loading amount	Scan rate mV s <sup>-1</sup>	Overpotential at specific current density	Tafel slope mV dec <sup>-1</sup>	Reference
Defective MnCo <sub>2</sub> O <sub>4</sub>	0.1 M KOH	~0.40 mg cm <sup>-2</sup>	10	400 mV at 10 mA cm <sup>-2</sup>	87	79
NiMoO <sub>4</sub>	1 M KOH	0.282 mg cm <sup>-2</sup>	5	340 mV at 10 mA cm <sup>-2</sup>	45.6	80
NiFeMo oxide	0.1 M KOH	0.25 mg cm <sup>-2</sup>	0.5	280 mV at 10 mA cm <sup>-2</sup>	49	81
NiCo <sub>2</sub> O <sub>4</sub>	1 M KOH	—	—	320 at 10 mA cm <sup>-2</sup>	30	82
NiCo <sub>2</sub> O <sub>4</sub>	1 M KOH	0.25 mg cm <sup>-2</sup>	10	—	47	68
NiFe <sub>2</sub> O <sub>4</sub> /CoFe <sub>2</sub> O <sub>4</sub>	1 M KOH	5 mg cm <sup>-2</sup>	5	240 mV at 10 mA cm <sup>-2</sup>	42	83
Ni <sub>0.67</sub> Fe <sub>0.33</sub> LDH/C	1 M KOH	0.28 mg cm <sup>-2</sup>	—	210 mV at 10 mA cm <sup>-2</sup>	35.1	84
Ni-Fe hydroxide	1 M KOH	32 μg cm <sup>-2</sup>	5	270 mV at 80 mA cm <sup>-2</sup>	28	85
NiCo <sub>2.7</sub> (OH) <sub>x</sub>	1 M KOH	0.14 mg cm <sup>-2</sup>	5	350 mV at 10 mA cm <sup>-2</sup>	65	86
NiVr-LDH	1 M KOH	—	—	180 mV at 10 mA cm <sup>-2</sup>	28	46
Zn <sub>x</sub> Co <sub>3-x</sub> O <sub>4</sub> nanoarray	1 M KOH	≈ 1 mg cm <sup>-2</sup>	0.5	320 mV at 10 mA cm <sup>-2</sup>	51	87
Spindle-like ZnCo <sub>2</sub> O <sub>4</sub>	1 M KOH (pH = ~14)	0.24 mg cm <sup>-2</sup>	5	389 mV at 10 mA cm <sup>-2</sup>	59.5	88
ZnCo <sub>2</sub> O <sub>4</sub>	0.1 M KOH	51 μg cm <sup>-2</sup>	—	420 mV at 10 mA cm <sup>-2</sup>	38.5	37
Cu <sub>0.7</sub> Co <sub>2.3</sub> O <sub>4</sub>	1 M KOH	100 μg cm <sup>-2</sup>	20	367 mV at 10 mA cm <sup>-2</sup>	59.6	89
SnCoFe hydroxide	1 M KOH	—	2	270 mV at 10 mA cm <sup>-2</sup>	42.3	90
CoO/Co <sub>3</sub> O <sub>4</sub>	1 M KOH	0.019 mg cm <sup>-2</sup>	5	260 at 10 mA cm <sup>-2</sup>	54	47
CoO-MoO <sub>2</sub>	1 M KOH	0.5 mg cm <sup>-2</sup>	5	312 mV at 10 mA cm <sup>-2</sup>	69	91
Co <sub>2</sub> Mo <sub>3</sub> O <sub>8</sub> @NC	1 M KOH	0.14 mg cm <sup>-2</sup>	5	331 mV at 10 mA cm <sup>-2</sup>	87.5	92
CoVO <sub>x</sub>	1 M KOH	0.5 mg cm <sup>-2</sup>	10	330 mV at 10 mA cm <sup>-2</sup>	46	93
CoFe <sub>2</sub> O <sub>4</sub>	0.1 M KOH	0.15 mg cm <sup>-2</sup>	10	412 mV at 10 mA cm <sup>-2</sup>	53	94
Co-MoO <sub>x</sub>	1 M KOH	0.28 mg cm <sup>-2</sup>	—	340 mV at 10 mA cm <sup>-2</sup>	49	95
CoO <sub>x</sub> (Ce)	1 M KOH	1.1 mg cm <sup>-2</sup>	—	261 mV at 20 mA cm <sup>-2</sup>	65.67	96
LaMn <sub>7</sub> O <sub>12</sub>	0.1 M KOH	0.25 mg cm <sup>-2</sup>	10	—	~100	97
LaSr <sub>1.4</sub> Ni <sub>0.6</sub> MnO	1 M KOH	0.25 mg cm <sup>-2</sup>	10	367 mV at 1 mA cm <sup>-2</sup>	86	98
LaCoO <sub>3</sub> -80 nm	0.1 M KOH	0.25 mg cm <sup>-2</sup>	—	490 mV at 10 mA cm <sup>-2</sup>	69	57
La(Co <sub>0.71</sub> Ni <sub>0.25</sub> ) <sub>0.96</sub> O <sub>3-δ</sub>	0.1 M KOH	0.7 mg cm <sup>-2</sup>	10	324 mV at 10 mA cm <sup>-2</sup>	—	99
La <sub>0.2</sub> Sr <sub>0.8</sub> FeO <sub>3</sub>	0.1 M KOH	0.25 mg cm <sup>-2</sup>	10	318 mV at 50 μA cm <sup>-2</sup>	58	63
La <sub>0.2</sub> Sr <sub>0.8</sub> FeO <sub>3</sub>	0.1 M KOH	—	10	430 mV 34.6 μA cm <sup>-2</sup>	71	100
La <sub>0.5</sub> Sr <sub>1.5</sub> Ni <sub>0.7</sub> Fe <sub>0.3</sub> O <sub>4.04</sub>	0.1 M KOH	0.015 mg cm <sup>-2</sup>	—	360 mV at 10 mA cm <sup>-2</sup>	—	101
La <sub>0.5</sub> Ba <sub>0.25</sub> Sr <sub>0.25</sub> CoO <sub>2.9-δ</sub> F <sub>0.1</sub>	1 M KOH	0.157 mg cm <sup>-2</sup>	—	~518 mV at 100 mA cm <sup>-2</sup>	113	45
LiCo <sub>0.8</sub> Fe <sub>0.2</sub> O <sub>2</sub>	0.1 M KOH	0.232 mg cm <sup>-2</sup>	5	340 mV at 10 mA cm <sup>-2</sup>	50	102
Li <sub>0.5</sub> Ni <sub>0.5</sub> O	1 M KOH	0.3 mg cm <sup>-2</sup>	10	~	56.6	103
SrCo <sub>1-x</sub> Ru <sub>x</sub> O <sub>3-δ</sub>	0.1 M KOH	0.23 mg cm <sup>-2</sup>	5	360 mV at 10 mA cm <sup>-2</sup>	113	104
SrCo <sub>0.95</sub> Po <sub>0.05</sub> O <sub>3-δ</sub>	0.1 M KOH	0.232 mg cm <sup>-2</sup>	10	480 mV at 10 mA cm <sup>-2</sup>	84	105
BaNi <sub>0.83</sub> O <sub>2.5</sub>	0.1 M KOH	0.295 mg cm <sup>-2</sup>	10	~460 mV at 10 mA cm <sup>-2</sup>	—	106
Ba <sub>4</sub> Sr <sub>4</sub> (Co <sub>0.8</sub> Fe <sub>0.2</sub> ) <sub>4</sub> O <sub>15</sub>	0.1 M KOH	0.232 mg cm <sup>-2</sup>	10	340 mV at 10 mA cm <sup>-2</sup>	47	107
BaCo <sub>0.7</sub> Fe <sub>0.2</sub> Sn <sub>0.1</sub> O <sub>3-δ</sub>	0.1 M KOH	0.232 mg cm <sup>-2</sup>	10	~420 mV at 10 mA cm <sup>-2</sup>	69	108
PrBa <sub>0.5</sub> Sr <sub>0.5</sub> Co <sub>1.5</sub> Fe <sub>0.5</sub> O <sub>5+δ</sub>	0.1 M KOH	0.202 mg cm <sup>-2</sup>	—	360 mV at 10 mA cm <sup>-2</sup>	55	109
Nd <sub>1.5</sub> Ba <sub>1.5</sub> CoFeMnO <sub>9-δ</sub>	0.1 M KOH	0.209 mg cm <sup>-2</sup>	10	359 mV at 10 mA cm <sup>-2</sup>	81	110
Fe <sub>2</sub> O <sub>3</sub>	0.5 M H <sub>2</sub> SO <sub>4</sub>	—	5	650 mV at 10 mA cm <sup>-2</sup>	56	111
Cu <sub>1.5</sub> Mn <sub>1.5</sub> O <sub>4</sub> :10F	0.5 M H <sub>2</sub> SO <sub>4</sub>	1 mg cm <sup>-2</sup>	5	320 mV at 9.15 mA cm <sup>-2</sup>	60	112
Co <sub>3</sub> O <sub>4</sub>	0.5 M H <sub>2</sub> SO <sub>4</sub>	—	10	370 mV at 10 mA cm <sup>-2</sup>	70	113
Co <sub>0.05</sub> Fe <sub>0.95</sub> O <sub>y</sub>	0.05 M H <sub>2</sub> SO <sub>4</sub> (pH = 0.3)	—	—	650 mV at 10 mA cm <sup>-2</sup>	110	114
Ti-MnO <sub>2</sub>	0.05 M H <sub>2</sub> SO <sub>4</sub> (pH = 0.3)	—	—	>670 mV at 10 mA cm <sup>-2</sup>	170	115
MnO <sub>x</sub>	In phosphate buffer pH = 7	—	20	—	~70	116
Mn <sub>0.67</sub> Sb <sub>0.33</sub> O <sub>z</sub>	1 M H <sub>2</sub> SO <sub>4</sub> (pH = 0.13)	—	10	580 mV at 50 mA cm <sup>-2</sup>	75	117
Ni <sub>0.2</sub> Mn <sub>0.8</sub> Sb <sub>1.6</sub> O <sub>y</sub>	1 M H <sub>2</sub> SO <sub>4</sub>	—	—	—	~60	118
NiFePbO <sub>x</sub>	0.1 M KP <sub>i</sub> + 1 M KNO <sub>3</sub> (pH = 2.5)	—	—	>700 mV at 1 mA cm <sup>-2</sup>	90	119
CaCu <sub>3</sub> Ru <sub>4</sub> O <sub>12</sub>	0.5 M H <sub>2</sub> SO <sub>4</sub>	0.122 mg cm <sup>-2</sup>	—	—	40	120

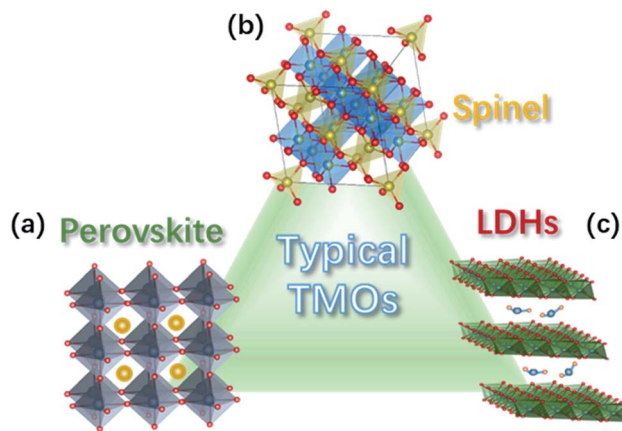


Fig. 2 Crystal structure of common TM oxides for OER: (a) perovskite, (b) spinel, and (c) LDH structure.

descriptors could help us rationally design TM oxide catalysts with improved performance for the OER and push forward for further developments in this field.

In this review, we provide an overview of the recent progress of the researches on OER catalysts and the generally proposed reaction mechanisms along with the emphasis on the correlation between oxide electronic structures and OER activity. We firstly introduce the most common TM oxide electrocatalysts studied for alkaline water electrolysis along with some recent reports of outstanding catalytic performances, and the main reaction mechanisms that have been proposed for the OER on TM oxides. Then, the electronic structure descriptors in TM oxides for the OER electrochemical activity are discussed, followed by a review on material science strategies to tailor the electronic structure parameters for the design of highly active electrocatalysts. Finally, we briefly conclude the review and discuss the challenges and opportunities for the future development of TM oxides. The in-depth comprehension of the correlation between the OER activity and electronic structure of TM oxide catalysts can be instructive to explore future generation catalysts, and connects the research fields of heterogeneous catalysis with the solid-state chemistry and physics.

## 2. TM oxide electrocatalysts for alkaline OER

A major advantage of the anion-exchange membrane water electrolysis over the proton-exchange membrane electrolysis is the replacement of conventional noble metal electrocatalysts, such as the traditional Ir- and Ru-based oxides, with low-cost TM oxides. Additionally, the alkaline media provides higher stability to the OER metal oxide catalysts. Since the TMs in the oxide have significant contribution to the valence states of the material, the physicochemical properties are strongly susceptible to the type and chemistry (*e.g.*, oxidation state, coordination and spin state) of the TM. This property provides TM oxides with large flexibility for the design of electrocatalysts for the OER. As a result, many metal oxides have been suggested as

promising OER catalysts in alkaline media over the last decade.<sup>20,60</sup> Table 1 provides a comparison of the performance of various TM oxides electrocatalysts for the OER recently reported in the literature, which provides context to the current state of the art. Although TM oxides in a wide range of compositions are currently under investigation, most can be grouped into complex perovskite oxides, spinel oxides or layered (oxy)hydroxides.

### 2.1 Perovskite TM oxides

Perovskite TM oxides are attractive electrocatalysts for promoting the OER because of their tuneable electronic structure and high activity.<sup>61</sup> Perovskite oxides have a general  $ABO_3$  structure, with rare-earth and/or alkaline large cations at the A site that have a 12-fold coordination with oxygen anions, and TMs at the B site in the centre of a corner-sharing octahedra of oxygen anions, see Fig. 2a. The chemistry of the TM in the B site can be widely modulated by substitution at A site, *e.g.*,  $La^{3+}$  for  $Sr^{2+}$ , which provides means to tailor the OER catalytic properties of the oxides.<sup>51,61</sup> Well-known examples of this approach are the systems  $Sr_xLa_{1-x}FeO_3$  and  $Sr_xLa_{1-x}CoO_3$  with  $0 \leq x \leq 1$ , in which the substitution of  $La^{3+}$  for  $Sr^{2+}$  induces the oxidation of the TM at the B site, leading to improvements of the catalytic activity for the OER.<sup>54,62,63</sup>

### 2.2 Spinel oxides

Spinel oxides have also been shown to drive the OER with excellent catalytic performances.<sup>64</sup> These oxides have the general chemical formula  $AB_2O_4$  and a crystalline structure in which TM cations in tetrahedral (cation A) and octahedral sites (cation B) coexists, as depicted in Fig. 2b. Such arrangement provides multiple sites for accommodating different transition-metal cations and in a wide range of valence states, allowing therefore the preparation of a wide variety of oxides.<sup>64,65</sup> The possibility for TM cations in both the tetrahedral and octahedral sites to be involved in the OER may bring ambiguities as to which site is responsible for the OER catalysis. However, theoretical and experimental studies have specifically addressed this issue, revealing that the octahedral sites have a superior OER catalytic activity, therefore, those are generally considered the active catalytic sites for the OER.<sup>37,66</sup> The chemistry of the TM in the octahedral site can be tuned by the synthesis conditions and the composition, which provides means to tailor the catalytic activity for the OER.<sup>67,68</sup> Ni in  $Ni_xCo_{3-x}O_4$  can be prepared in the  $0 \leq x \leq 1$  range,<sup>68</sup> whereas  $Mn_xCo_{3-x}O_4$  can be obtained as a single phase in the whole  $0 \leq x \leq 3$  range.<sup>67</sup> In both cases, the reorganization of the cation site and their oxidation states leads to substantial modulation of the catalytic activity of the OER.

### 2.3 TM oxyhydroxides

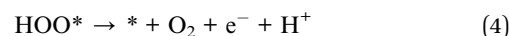
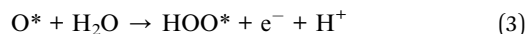
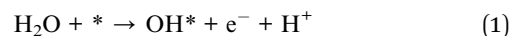
TM oxyhydroxides formed on the surface of TM have been long recognised as highly active catalytic species for OER in alkaline medium.<sup>69</sup> Ni oxyhydroxide have received extensive attention because of its particularly high catalytic activity;<sup>70</sup> however, the observed high performance have been shown to arise from the presence of Fe impurities.<sup>71,72</sup> In a detailed study, L. Trotochaud

*et al.* found that the incorporation of Fe impurities in Ni-based OER electrodes leads to the formation of a mixed Ni-Fe layered double hydroxides (NiFe-LDH) with a fougèrite/pyroaurite mineral structure (Fig. 2c), which is responsible of the high catalytic activity.<sup>73</sup> In this structure, both Ni and Fe cations reside in octahedral local coordination; however, it was also shown that the long-range order is not important for the high activity of NiFe-LDH. Optimised synthesis of NiFe-LDH electrodes have led to overpotentials as low as 130 mV to achieve a current density of 10 mA cm<sup>-2</sup>,<sup>74</sup> which has motivated a large number of studies on this type of OER catalysts, in recent years.<sup>75</sup> The possibility of intercalating the LDH structure increases the flexibility for material design. For instance, the catalytic activity for the OER of NiFe LDHs can be tailored by the intercalation of anions, and it has been shown that the activity of the intercalated NiFe-LDH varies with the redox potential of the anion: low redox potential of hypophosphites favours the catalysis of the OER.<sup>76</sup> Similar to complex oxides, the substitutional doping of NiFe LDHs is also used for tuning the catalytic activity. For example, V-doping has led to an outstanding catalytic performance due to enhanced conductivity, facile electron transfer, and abundant active sites.<sup>77</sup> Dionigi and co-workers provided direct evidence that, under applied anodic potentials, NiFe and other Fe-containing LDHs oxidise from as-prepared  $\alpha$ -phases to activated  $\gamma$ -phases. The OER-active  $\gamma$ -phases are characterised by about 8% contraction of the lattice spacing and switching of the intercalated ions.<sup>78</sup>

### 3. OER mechanisms for TM oxide electrocatalysts

Early attempts to describe the OER with parameters that could explain and predict the catalytic activity of TM oxides were based on the bond strength between surface of metal oxides and the adsorbates. Based on kinetic studies, Bockris and Otagawa the concluded that the OER on a series of perovskite oxides, containing metal cations from the first transition row, consisted in the electrochemical adsorption of OH<sup>-</sup>, followed by the rate-determining electrochemical desorption of  $\cdot$ OH radical, which

reacts with water to form O<sub>2</sub>.<sup>50,121</sup> The study of adsorbates bonding strengths as descriptors of the OER have been successfully developed with the aid of *ab initio* density functional theory (DFT) calculations, resulting in the elucidation of a well-established adsorbate evolution mechanism (AEM) depicted below in eqn (1)–(4).<sup>122–125</sup>



Rossmeisl and co-workers investigated trends in the electrocatalytic OER on basis of a database of O\*, HO\* and HOO\* adsorption energies on oxide surfaces, finding that there is a linear scaling relationship between the adsorption energies of the three species.<sup>123</sup> In particular, the adsorption energies of HO\* and HOO\* are closely related because both species bind with the catalyst surface through an oxygen atom with a single bond so that the difference in adsorption energy for these two species ( $\Delta G_{\text{HOO}^*}^0 - \Delta G_{\text{HO}^*}^0$ ) is constant for any catalyst surface (see Fig. 3a).<sup>126</sup> Thus, the variation in the OER overpotential between different oxide surfaces is determined by the O\* adsorption energy. For catalysts that bind oxygen adsorbates strongly to its surface, the formation of the O–O bond in the HOO\* adsorbate will be slow, making step 3 (eqn (3)) the rate-determining step (RDS). On the other hand, for catalysts that form labile bonds with oxygen adsorbates the deprotonation of the HO\* species will be slow, making step 2 (eqn (2)) the RDS. Another consequence of the fixed adsorption energy difference between adsorbates HO\* and HOO\* is that the  $\eta^{\text{OER}}$  can be described in terms of the formation free energies of O\* and HO\*,  $\Delta G_{\text{O}^*}^0$  and  $\Delta G_{\text{HO}^*}^0$ , respectively. Plotting  $\eta^{\text{OER}}$  vs. ( $\Delta G_{\text{O}^*}^0 - \Delta G_{\text{HO}^*}^0$ ) leads to a general volcano-type relationship independent of the catalysts, as shown in Fig. 3b and c. Therefore, in order to rationalised the OER activity, attention must be focused on the oxygen binding strength on the surface of catalysts

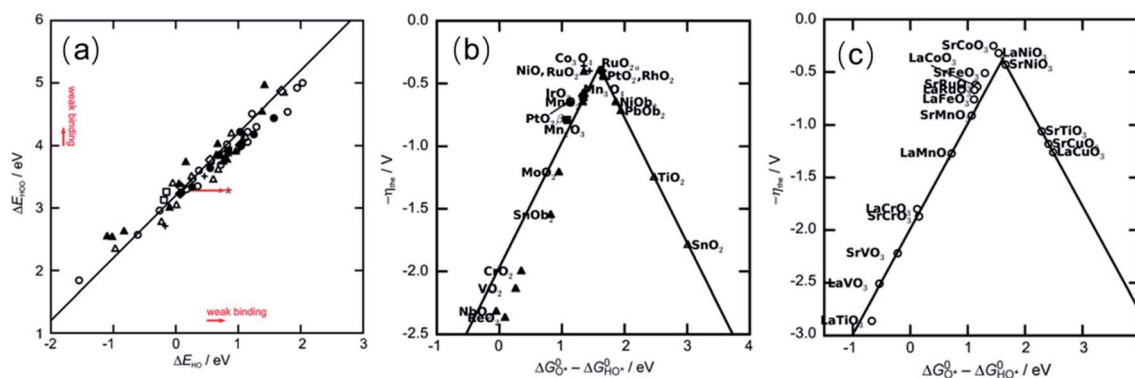


Fig. 3 (a) Adsorption energy of HOO\* plotted against the adsorption energy of HO\* on perovskite, rutile, anatase, Mn<sub>x</sub>O<sub>y</sub>, Co<sub>3</sub>O<sub>4</sub>, and NiO oxides. (b) Trends for the OER activity on rutile, anatase, Mn<sub>x</sub>O<sub>y</sub>, Co<sub>3</sub>O<sub>4</sub>, and NiO oxides. (c) Trends for the OER activity on perovskite. Reproduced with permission.<sup>123</sup> Copyright 2011, Wiley-VCH.

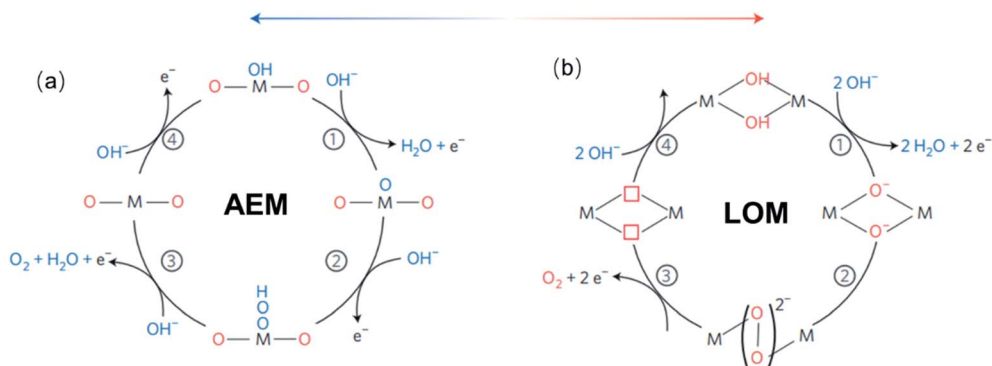


Fig. 4 (a) The classical four-steps OER mechanism of proton-coupled electron transfer (AEM), (b) the idealized OER mechanism based on the direct pairing of two lattice oxygens (LOM) in the anionic redox process. Reproduced with permission.<sup>127</sup> Copyright 2016, Macmillan Publishers Limited.

following the Sabatier principle: the most active catalyst binds oxygen intermediate species to the surface neither too strongly nor too weakly. The most active catalysts are in the centre of this plot and bind these key reaction intermediates neither too strongly nor weakly.

For TM oxides with high covalency in the TM–O bonds, the adsorption energy of oxygen adsorbates generally decreases and the direct O–O bonding with the reversible formation of oxygen vacancy become favourable. In this scenario, a second OER mechanism, the so-called lattice oxygen mechanism (LOM), has been described.<sup>126,127</sup> Different from the AEM, the LOM proceeds on two neighbouring metal sites, and consists of 4 steps: (1) two HO\* on the metal sites undergo deprotonation, resulting in two neighbouring TM–O<sup>−</sup> species, (2) these two oxo species directly couple to form the O–O bond, (3) O<sub>2</sub> is released leaving behind two oxygen vacancies, (4) reoccupation of these vacancies with OH<sup>−</sup>. Fig. 4 shows a graphic comparison of the AEM and the LOM. Since HOO\* is not an intermediate in the LOM catalytic cycle, catalysts that drive the OER *via* this mechanism deviate from the trends derived from the AEM which are based on the scaling relation between HO\* and HOO\*. Over the past few years, the LOM has been confirmed by a series of observations that demonstrate the importance of the bulk oxide electronic structure in the catalytic OER,<sup>128,129</sup> and direct evidences of involvement of lattice oxygen in the OER.<sup>130–132</sup>

## 4. Electronic structure parameters in TM oxides for OER electrochemical activity

From the reaction mechanisms (AEM and LOM) described in the previous section, it is clear that the OER process is driven by (i) *the binding strength of the intermediate oxygen adsorbates (O\* and OH\*)*, which is important when the AEM operates, and (ii) *the lattice TM–O covalency*, which dictates when the LOM becomes active. Additionally, electrons transfer between the electrolyte and the catalyst's surface will always be involved in the OER, and therefore (iii) *the Schottky barrier* for such electrons transfer,<sup>133</sup> is also an important factor the overall OER

process. All of these three factors are derived from the electronic structures of the TM oxides, which provides means to directly correlate measurable and controllable electronic properties with the catalytic activity. On the one hand, the electronic configuration of the TM cations drives the covalency of the TM–oxygen bond in the bulk structure, which is ultimately related to the surface TM binding energies with adsorbate oxide species. A coordination chemistry approach is useful for describing these bonds.<sup>61</sup> The local structure of most TM metal oxides catalytically active for the OER, *e.g.*, perovskite, rutile, spinel and LDH, feature a TM bonded to six O atoms in an octahedral (TM)O<sub>6</sub> coordination. In such coordination, the TM 3d and oxygen 2p atomic orbitals hybridize to form  $\sigma$  orbitals (O 2p with  $d_{z^2}$  and  $d_{x^2-y^2}$  orbitals of the TM) and  $\pi$  orbitals (O 2p orbitals with  $d_{xy}$ ,  $d_{xz}$  and  $d_{yz}$  orbitals of the TM), as depicted in Fig. 5a. The  $\sigma$  and  $\pi$  antibonding ( $\sigma^*$  and  $\pi^*$ ) orbitals are named after their symmetry  $e_g$  and  $t_{2g}$ , respectively, and they are the orbital of interest because their filling determines the covalency of the TM–oxygen bonds. On the other hand, the Schottky barrier for an electron transfer between adsorbates (O\* and OH\*) and TM oxide catalysts is defined by the energy difference between the unoccupied TM 3d–O 2p band and the redox potential of the electrolyte.

On basis of those factors, several electronic structure descriptors for the OER activity have been recently proposed, addressing the binding strength of adsorbed OER intermediates, the covalency of the TM–O bond, and the energetics for electrons transfer at the interface. The most successful descriptors of the binding strength of adsorbates is the number of electrons in the  $e_g$ -orbital of the transition metal cations.<sup>51</sup> As for the TM-covalency, the degree of TM 3d–O 2p hybridization,<sup>128,129</sup> the energy position of the O 2p band centre,<sup>59</sup> and the charge-transfer energy, have been used as electronic descriptors.<sup>35</sup> Whereas, the Schottky barrier for electrons transfer at the interface has been discussed in terms of the band alignment at the catalysts/electrolyte interface, using parameters like the work function and the electron affinity of the oxide relative to the electrolyte.<sup>35,63,68</sup> *In the following, we will provide the fundamental basis of these electronic descriptor, and how they are used in order to rationalise OER activity and mechanism. In the following Section 5, we*

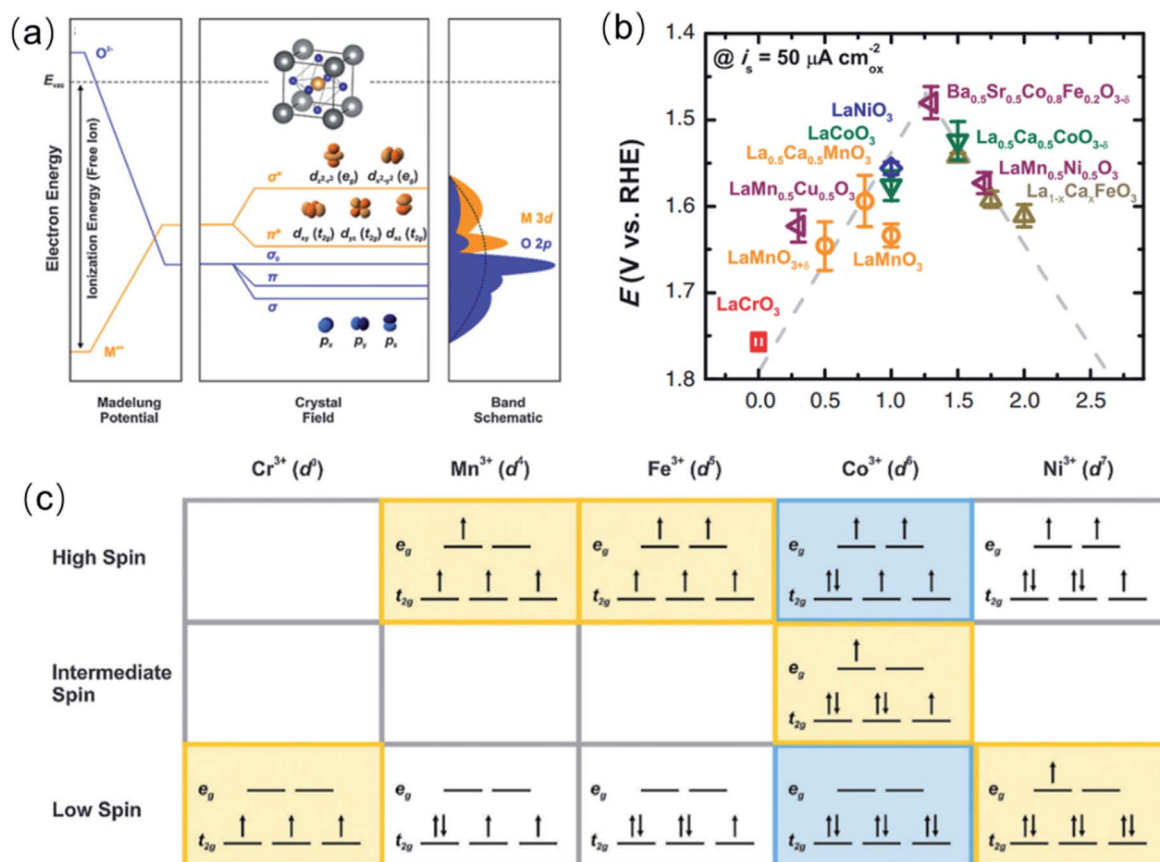


Fig. 5 (a) Schematic band structure of perovskite oxides (unit cell inset). (b) Elation between OER activity and  $e_g$  occupancy of transition metal cations (B in  $\text{ABO}_3$ ). (b) Is reproduced with permission.<sup>51</sup> Copyright 2011, American Association for the Advancement of Science. (c)  $e_g$  orbital occupancy of some transition metal cations with different spin states. (a and c) Are reproduced with permission.<sup>25</sup> Copyright 2015, Royal Society of Chemistry.

also discuss the strategies for modulating the electronic structure parameter in order to optimise the activity for the OER.

#### 4.1 Descriptors for the binding strength of the intermediate oxygen adsorbates: the number of electrons in the $e_g$ -orbitals of the TM cations

The filling of the  $e_g$ -orbitals as a descriptor of the OER catalytic activity of perovskite oxides was first proposed by Shao-Horn and co-workers.<sup>51</sup> Since the  $e_g$  orbitals of surface transition metal ions participate in  $\sigma$ -bonding with adsorbates, its occupancy influences the binding energy of oxygen-related intermediate species on the TM, and therefore the OER activity. They reported that a near-unity occupancy of the  $e_g$  orbitals of surface transition metal ions and high covalency in bonding to oxygen can enhance the intrinsic OER activity of perovskite transition metal oxides in alkaline solution. Specifically, they found that an occupancy of 1.2 in  $\text{Ba}_{0.5}\text{Sr}_{0.5}\text{Co}_{0.8}\text{Fe}_{0.2}\text{O}_{3-\delta}$  was optimum within a group of perovskite oxides with number of  $e_g$  electrons varying from 0 to 2 (Fig. 5b). The occupancy of  $e_g$ -orbitals is more relevant than that of  $t_{2g}$  orbitals because the  $\sigma$ -bonding  $e_g$  orbital has a stronger overlap with the oxygen-based adsorbate than does the  $\pi$ -bonding  $t_{2g}$  orbital, and therefore can more

directly promote electron transfer between surface cation and adsorbed reaction intermediates.<sup>51</sup> More recently, it was found that the  $e_g$  occupancy of the cation in the octahedral site of spinel oxides also works as a descriptor of the OER catalytic activity of these materials,<sup>67</sup> with a similar near unity optimum occupancy.

In general, the  $e_g$  orbital occupation of TM oxides of the first transition series can be modified by tuning the oxidation state and spin state of the TM cation. Although the total number of 3d electrons is only determined by the oxidation state, e.g.,  $\text{Ni}^{3+}$  is a  $3d^7$  cation and  $\text{Fe}^{3+}$  is a  $3d^5$  cation, the spin configuration changes the  $e_g$ -orbital occupancy. As shown in Fig. 5c,  $\text{Ni}^{3+}$  can adopt a low spin configuration  $t_{2g}^6e_g^1$  or a high spin configuration  $t_{2g}^5e_g^2$ , respectively providing 1 or 2  $e_g$  electrons, whereas  $\text{Fe}^{3+}$  can have 2  $e_g$  electrons in the high spin configuration  $t_{2g}^3e_g^2$  or none in the low spin configuration  $t_{2g}^5e_g^0$ . Therefore, characterisation techniques for the determination of the actual oxidation state and spin configuration are very important in order to rationalise the catalytic activity for the OER based on the  $e_g$  occupancy. X-ray photoelectron spectroscopy (XPS) and near-edge X-ray absorption fine structure (NEXAFS) are generally used to explore the oxidation state of elements in a compound,<sup>134–136</sup> and most studies of OER catalysts include such spectroscopic



characterization. Although the characterisation of the spin configuration is less extended, some authors have specifically studied the magnetic properties and spin configurations of TM oxide OER catalysts using field-cooling magnetization measurements,<sup>58,137</sup> and superconducting quantum design (SQUID) magnetometer.<sup>138</sup> From these measurements, an effective magnetic moment  $\mu_{\text{eff}}$  can be calculated, which provides means to determine the average spin configuration. The study of magnetic properties has been particularly extensive in perovskite cobaltite because whether the spin state of  $\text{Co}^{3+}$  ions at room temperature is a mixture of low spin ( $t_{2g}^6e_g^0$ ) and high spin ( $t_{2g}^4e_g^2$ ) states or an intermediate-spin ( $t_{2g}^5e_g^1$ ) state (see Fig. 5c) has been controversial in the past decades. However, using either interpretation of the spin state in perovskite cobaltite, there is a general agreement that an average  $e_g$  occupancy close to the unity is beneficial for the OER activity. Advanced interpretations of the photoelectron spectra and the X-ray adsorption spectra also provide the magnetic properties;<sup>139</sup> however well-defined samples are usually required.

**4.1.1 OER mechanism based on the  $e_g$ -orbitals occupancy.** The  $e_g$ -orbitals occupancy is closely related to the AEM because the binding strength of adsorbate oxygen species depends on the  $e_g$  orbitals occupancy. A low occupancy, in the left branch of the volcano plot, favours a strong bond oxygen species so that the formation of the O–O bond in the  $\text{HOO}^*$  adsorbate (step 3 of the AEM depicted in eqn (3)) is the rate-determining step (RDS). On the other hand, a high occupancy, the right branch of the volcano plot, favours the covalent more labile oxygen bonds that may retard the deprotonation of the  $\text{HO}^*$  species (step 2 of the AEM depicted in eqn (2)), making it the RDS. According to the  $e_g$  filling descriptor, an occupancy close to unity balances the strength of the oxygen adsorbates, efficiently facilitating these RDS, which leads to the highest OER activity.<sup>51</sup> The  $e_g$ -occupancy descriptor is based on the AEM ionic model, which is focused on the oxygen binding strength on the surfaces of the catalysts. Consequently, the predicted best OER catalysts are those binding oxygen species not too strong nor too weakly, following the Sabatier principle. However, this descriptor alone fails to explain why, for instance, the catalytic activity of



**Fig. 6** (a) Shows the potentials to reach  $50 \mu\text{A cm}^{-2}$  current density as a function of the  $H$  for  $\text{LaMnO}_3$ ,  $\text{LaCoO}_3$  and  $\text{LaNiO}_3$ . Reproduced with permission.<sup>51</sup> Copyright 2011, American Association for the Advancement of Science. (b) Schematic representation of the O  $p$ -band for transition metal oxides and evolution of the  $iR$ -corrected potential at  $0.5 \text{ mA cm}^{-2}$  oxide versus the O  $p$ -band centre relative to  $E_F$  (eV). Reproduced with permission.<sup>60</sup> Copyright 2013, Macmillan Publishers Limited. (c) Representations of the transition metal and oxygen partial density of states based on the ionic model, density functional theory and X-ray spectroscopy (XES, XAS, and XPS) of  $\text{LaCoO}_3$  and correlation between charge transfer energy ( $\Delta$ ) and the OER activity of 10 metal oxides, 5 semi-metallic (open dots) and 5 semiconducting (closed dots). Reproduced with permission.<sup>134</sup> Copyright 2017, Royal Society of Chemistry.

LaMnO<sub>3</sub> is much lower than those of LaCoO<sub>3</sub> and LaNiO<sub>3</sub>, when all these perovskite oxides have the same  $e_g$  orbital occupancy ( $e_g = 1$ ). This is a result of the ionic model failing to describe the covalency of the TM–O bond and the participation of lattice oxygen in the catalytic OER, which have been shown to play an important role in highly covalent late TM oxide catalysts.<sup>130–132</sup>

#### 4.2 Descriptors of the TM–O covalency: degree of TM 3d–O 2p hybridisation

Aiming at a more comprehensive description of the catalytic OER on complex perovskites, Shao-Horn and co-workers introduced a quantitative measure of the TM–O covalency as an additional descriptor for oxides with an  $e_g$ -occupancy of 1.<sup>51</sup> They estimated the TM 3d–O 2p hybridisation degree ( $H$ ) from the background-corrected spectral intensity associated with the O 1s  $\rightarrow$  TM 3d–O 2p transition within the O K-edge X-ray absorption (Abs) normalised by the nominal number TM 3d empty levels in  $e_g$  ( $h_{e_g}$ ) and  $t_{2g}$  ( $h_{t_{2g}}$ ) symmetries, using the expression:  $H = \text{Abs}/(h_{e_g} + \frac{1}{4}h_{t_{2g}})$ .<sup>51,140</sup>  $H$  is a measure of the covalency because the O K-edge XAS probes the transition from O 1s to unoccupied states with partial O 2p character hybridised with the valence states of the metal cations, due to dipole selection rules. Fig. 6a shows the potentials to reach 50  $\mu\text{A cm}^{-2}$  current density as a function of the  $H$  for LaMnO<sub>3</sub>, LaCoO<sub>3</sub> and LaNiO<sub>3</sub>. The linear decrease of the overpotentials with the increase of the TM–O covalency led the authors to conclude that a greater covalency between the TM and the oxygen, where the active redox couple of the TM cation has a larger O 2p character, promotes the charge transfer between surface cations and adsorbates, therefore accelerating the rate-determining steps in the OER.

Suntivich and co-workers estimated empirical trends of the TM 3d–O 2p hybridisation in complex perovskite using the pre-edge feature of the O K-edge X-ray absorption spectrum.<sup>141</sup> The authors found that the hybridisation increases with the number of 3d electrons, which is consistent with the expected electronegativity trend, and also with the oxidation state of the TM cation. It was recognised that a higher oxidation state of the TM in the oxide reduces the energy difference between the TM 3d and the O 2p, leading to an increased covalency.<sup>141,142</sup>

In 2013, Shao-Horn and co-workers demonstrated that the energy level of the O 2p band relative to the Fermi level ( $E_F$ ) could be also used as a descriptor for the TM–O covalency and activity of complex perovskite catalysts for the OER, as well as their stability.<sup>59</sup> Fig. 6b shows how the shift of the O 2p band centre closer to the  $E_F$  in a series of perovskite cobaltite, correlates with the decrease of the overpotential needed to reach a 0.5  $\text{mA cm}^{-2}$ . However, they found a shift limit from which the materials undergo amorphization during the OER, which leads to instability and activity losses. These observations suggested the participation of lattice oxygen in the reaction, which has been later confirmed for a series of perovskite cobaltite.<sup>130</sup> Although the level of the O 2p band relative to the  $E_F$  provides satisfying description of the OER on highly conducting oxides, is less reliable to describe OER kinetics on semiconductors and insulating oxides because the Fermi level may lie in forbidden energies and can be poorly defined for this type of materials.

More recently, Hong and co-workers made use of a combination of X-ray emission, X-ray absorption, and X-ray photoelectron spectroscopies to examine the electronic structure of a series of complex perovskite oxides LaBO<sub>3</sub> (B = Cr, Mn, Fe, Co and Ni).<sup>143</sup> These studies allowed the direct determination of the energy difference between the TM 3d and the O 2p, the so-called charge transfer energy ( $\Delta$ ), which can be directly related with the TM 3d–O 2p hybridisation degree, and therefore with the covalency in the TM–O bonds. The group of authors later reported that reducing  $\Delta$  greatly enhance the OER activity of oxides (Fig. 6c), and the trend across 10 oxides was statistically stronger than using the O 2p-band centre as the descriptor for both semi-metallic and semi-conducting oxides.<sup>35</sup> They showed that the charge-transfer energy, determined from the experimental elucidation of the band positions, can be used to evaluate trends in the electronic and chemical interactions at the oxide/electrolyte interface relevant to the OER, specifically the electron-transfer kinetics (electrochemical steps) and hydroxide affinity (chemical steps). As a result, the analysis of the charge-transfer energy of catalytic metal oxides can explain not only OER activities trends but also changes in the mechanisms of the OER from electron-transfer-limited to proton-electron-coupled, to proton-transfer-limited reactions.



Fig. 7 Schematic representation of the energy level versus density of states  $N(\epsilon)$ , showing the respective motion of the metal d band with respect to the oxygen p band in going from cationic to anionic redox processes and then O<sub>2</sub> release. Reproduced with permission.<sup>127</sup> Copyright 2016, Macmillan Publishers Limited.

**4.2.1 OER mechanism based on the degree of TM 3d–O 2p hybridisation.** On basis of the AEM, the enhanced OER activity observed upon increasing the covalent mixing in the TM–O bond at a constant  $e_g$  filling (Fig. 6a) was suggested to derive from a larger O 2p character of the active redox couple of the TM ion, which promotes the charge transfer between surface cations and key adsorbates, improving therefore the OER activity.<sup>51</sup> The beneficial effect of a high TM–O covalency on the OER catalytic activity have been shown to come together with substantial surface amorphization of the samples under the OER process,<sup>59,132,144</sup> even reaching an upper limit from which further increases in covalency lead to bulk amorphization and partial deactivation of the catalysts.<sup>59</sup> Such changes in surface morphology highlights the role of the TM–O covalency in the participation of lattice oxygen in the catalytic OER, which have been rationalised on basis of different electronic descriptors of the OER. As the TM–O covalency increases, for instance as  $x$  increases from 0 to 1 in systems like  $\text{Sr}_x\text{La}_{1-x}\text{CoO}_3$  or  $\text{Sr}_x\text{La}_{1-x}\text{FeO}_3$ , the Fermi level moves closer to the O 2p band centre (Fig. 7). As the Fermi level moves down in energy and closer to the O 2p states of the oxide, the antibonding TM 3d states below the Fermi level exhibit greater oxygen character, *i.e.*, there is a greater covalency of the metal–oxygen bond. The oxidation of lattice oxygen in the perovskites becomes thermodynamically favourable when O 2p states at the perovskite Fermi level lie above the redox energy of the  $\text{O}_2/\text{H}_2\text{O}$  couple, as shown in Fig. 7, triggering therefore the LOM.<sup>127</sup> Using *in situ*  $^{18}\text{O}$  isotope labelling mass spectrometry, Grimaud and co-workers provided direct experimental evidence that the  $\text{O}_2$  generated during the OER on some highly active oxides can come from lattice oxygen.<sup>130</sup> Density functional theory was employed by Yoo and co-workers to investigate the electronic origin and feasibility of the LOM on perovskites, finding that the different thermodynamic driving forces of lattice oxygen participation on  $\text{LaMO}_{3-\delta}$  ( $M = \text{Ni}, \text{Co}, \text{and Cu}$ ) are explained by the changes in the oxidation state of the TM site throughout the reaction.<sup>145</sup> Further DFT studies by this group of authors suggested that the LOM can lead to higher OER activity than the AEM by minimizing the thermodynamically required overpotential. Interestingly, they showed that whereas the OER activity volcano for the AEM is universal for all perovskites, that for the LOM depends on the identity of the A cation in  $\text{ABO}_3$ .<sup>146</sup>

*In situ*  $^{18}\text{O}$  isotope labelling mass spectrometry has been also employed to provide evidences of the LOM in oxides other than covalent perovskites. Amin and co-workers used isotope labelling together with differential electrochemical mass spectrometry to study the OER mechanism on spinel  $\text{Co}_3\text{O}_4$ .<sup>147,148</sup> Their work showed that the alkaline OER in  $\text{H}_2$   $^{18}\text{O}$ -containing electrolyte generated an amount of  $^{18}\text{O}^{16}\text{O}$  evolved increases from cycle to cycle before reaching a steady-state value, showing that a surface layer of  $\text{Co}_3\text{O}_4$  was taking part in OER *via* an oxygen exchange mechanism. They found that total number of oxygen atoms of the oxide participating in OER was 0.1 to 0.2% of the total oxide loading, corresponding to about 10–30% of the surface atoms, which represented the catalytically active sites.<sup>147</sup>

Although the LOM is associated with highly active covalent oxides, the OER mechanism for NiFe oxyhydroxides, the most

active OER catalyst in alkaline OER, has been shown not to proceed *via* lattice oxygen exchange.<sup>149,150</sup> Roy and co-workers used a model system of mass-selected NiFe nanoparticles and isotope labelling experiments to show that oxygen evolution in 1 M KOH does not proceed *via* lattice exchange on very highly active electrodes.<sup>150</sup> The lack of lattice oxygen participation in the OER on NiFe oxyhydroxides was further confirmed by Lee and co-workers in a mechanistic study based on  $^{18}\text{O}$ -labelling experiments in combination with *in situ* Raman.<sup>151</sup> However, in the same study, they also found that lattice oxygen does participate in the OER on Fe-free Ni and NiCo oxyhydroxides catalysts, specifically in the formation of a  $\text{NiOO}^-$  species that was identified as an OER intermediate. Based on DFT calculation, Huang and co-workers proposed that the OER proceeds *via* the LOM pathway on the metal oxyhydroxides only if two neighbouring oxidized oxygens can hybridize their oxygen holes without sacrificing metal–oxygen hybridization significantly, finding that the LOM is highly favoured in the OER on  $\text{Zn}_{0.2}\text{Co}_{0.8}\text{OOH}$ .<sup>152</sup>

### 4.3 Descriptor of the Schottky barrier for $e^-$ transfer: band alignment at the catalysts/electrolyte interface

As described in above in Section 4.2, the band energy position in metal oxides dictates the charge transfer energy, which is a measure of the covalency of the TM–O bond. Additionally, the band positions relative to the redox potential of the OER provides information on the electronic and chemical interaction at the catalysts/electrolyte interface relevant to the OER.<sup>35</sup> On the one hand, the energy difference between the unoccupied TM 3d–O 2p band and the OER redox potential defines the electron affinity of the oxide relative to the electrolyte, and it represents a Schottky barrier for electron transfer, as schematically shown in Fig. 8a. On the other hand, the energy difference between the potential of the electrolyte and the Fermi level is a measure of the hydroxide affinity of the catalyst (see Fig. 8a). In a comparative study of several oxides, Hong and co-workers found that oxides with a larger charge-transfer energy have a higher Schottky barrier for electron transfer at the oxide–electrolyte interface (Fig. 8b).<sup>35</sup> As the charge transfer energy decreases in the series of oxide catalysts, the barrier for electron transfer associated with OER decreases. Additionally, upon decreasing the charge transfer energy, the electrical conductivity of the oxides tends to increase, which brings about an extra benefit of a low resistance to the current flow during the OER.<sup>153</sup> However, the intrinsic OER active is more influenced by the presence of low-energy TM 3d orbital states that facilitates the interface electron transfer associated with the OER process. Yun and co-workers recently found that semiconducting  $\text{LaCoO}_3$  has an intrinsic OER activity more than one order of magnitude higher than that of metallic  $\text{LaNiO}_3$ , in contrast to previous reports.<sup>154</sup> The group of authors showed that misleading interpretations could be avoided by using very thin films of the catalyst on conducting interlayer as samples.

We have been recently working on a systematic approach to study the relationship between the OER catalytic activity and the electronic structure of complex perovskite, double perovskite and spinel systems such as  $\text{Sr}_x\text{La}_{1-x}\text{FeO}_3$ ,  $\text{La}_{2-x}\text{Sr}_x\text{NiMnO}_6$  and  $\text{Ni}_x\text{Co}_{3-x}\text{O}_4$ .<sup>63,68,98</sup> For all three systems, increasing  $x$  in the



Fig. 8 (a) Schematic representation of electron transfer energy at the electrolyte-oxide interface, hydroxide affinity of the oxide, and charge-transfer energy (covalency). (b) Estimation of electron transfer energy, hydroxide affinity, and charge-transfer energy from X-ray spectroscopic data under rigid band assumptions for La<sub>0.8</sub>Sr<sub>0.2</sub>CoO<sub>3-δ</sub> (LSCO20), La<sub>0.5</sub>Sr<sub>0.5</sub>CoO<sub>3-δ</sub> (LSCO50), GdBaCo<sub>2</sub>O<sub>5+δ</sub> (GBCO), SmBaCo<sub>2</sub>O<sub>5+δ</sub> (SBCO), PrBaCo<sub>2</sub>O<sub>5+δ</sub> (PBCO), Ba<sub>0.5</sub>Sr<sub>0.5</sub>Co<sub>0.8</sub>Fe<sub>0.2</sub>O<sub>3-δ</sub> (BSCF). Reproduced with permission.<sup>35</sup> Copyright 2017, Royal Society of Chemistry. (c) Left panel: the measured occupied and unoccupied density of states (DOS) near the E<sub>F</sub> for Co<sub>3</sub>O<sub>4</sub> and NiCo<sub>2</sub>O<sub>4</sub>; the energy level is relative to vacuum level (vs. Vac.). Right panel: schematics of the four-step reaction pathway of OER. Reproduced with permission.<sup>68</sup> Copyright 2019, American Chemical Society.

range  $0 \leq x \leq 1$  triggers a shift of the Fermi level ( $E_F$ ) closer to the valence band maximum (VBM) and induces the formation of hole states (unoccupied) below the conduction band minimum (CBM), near the  $E_F$ . We argue that *filled* electronic states (*i.e.*, VBM) near the  $E_F$  improve the orbital overlap with key intermediates, favouring their formation (Fig. 8d). At the same time, unoccupied states close to the  $E_F$  increase the electron affinity of the oxide semiconductors, thereby facilitating charge transfer at the interface associated with water oxidation. We found linear correlations between the OER activity and the downshift of the  $E_F$ , as well as with the hole state density, which shows that band energy level descriptors can be very powerful to describe the OER activity of TM oxides.

## 5. Material engineering strategies for the OER tailoring based on electronic parameters

As described above, in Section 4, many efforts have been devoted to discovering the correlation between the electronic structure and OER catalytic activity of TM oxide catalysts. Several electronic structure descriptors of important parameters for the OER catalysis, such as the adsorption energy of adsorbates (described by the  $e_g$  occupancy), the covalency (described by O 2p energy level and charge transfer energy) and the interface energetics (ionisation potential,  $E_F$  and electron

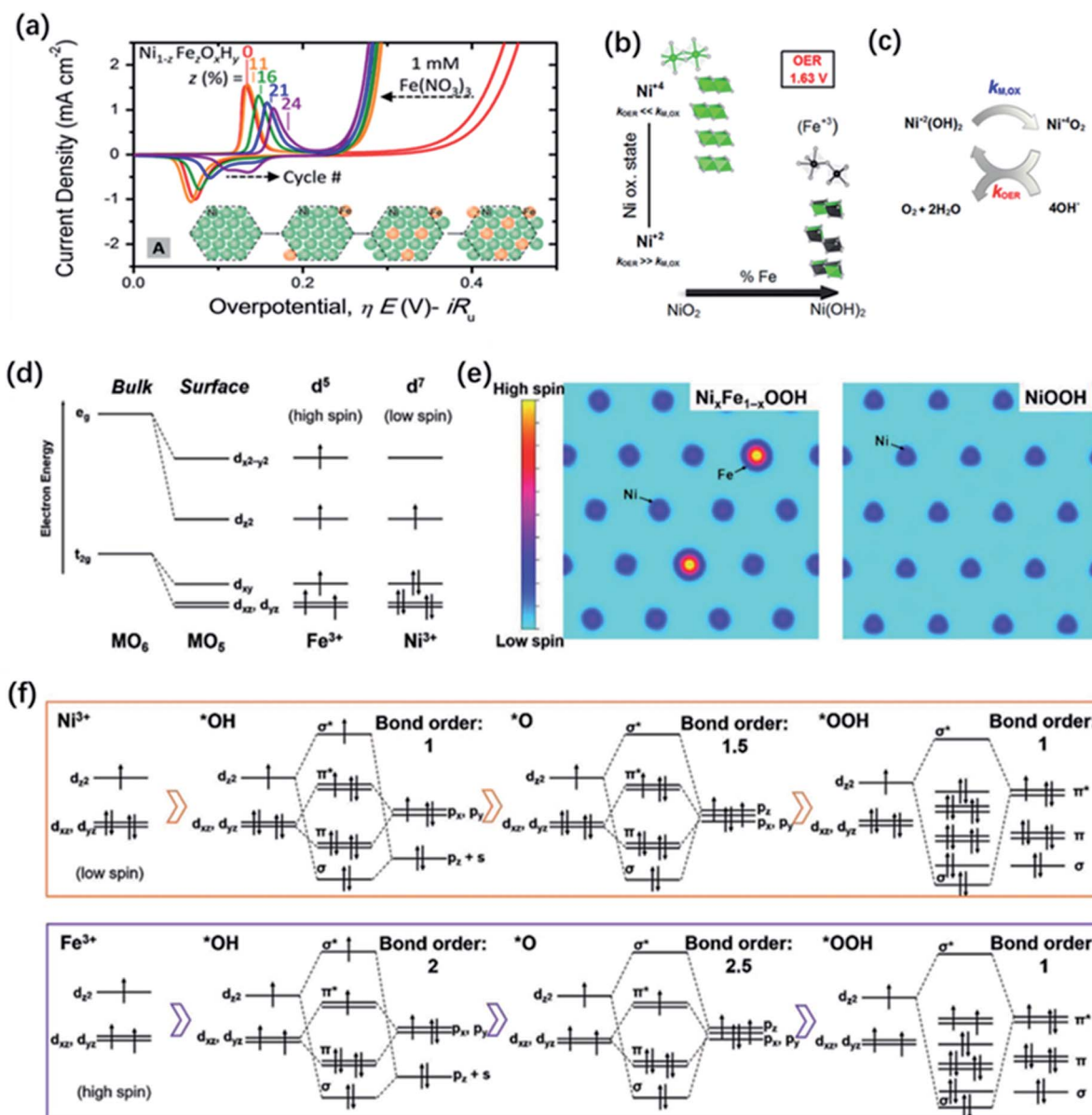
affinity), have been developed. In this section, we will focus on strategies to modify those electronic parameters that determine the OER catalytic properties of TM oxides through intentional doping, strain and defect engineering.<sup>155,156</sup>

### 5.1 Doping

Doping is a common and efficient strategy to tuning the electronic structure of TM oxide catalysts, due to the high flexibility in compositions and crystal structures. Introducing or substituting foreign metal ions with different radius size,

valence state and electron-donating ability may enable charge transfer occurred through shared oxygen ions or change the lattice parameters and thus the degree of overlap between atomic orbitals, finally altering the electronic, optical and magnetic properties of the oxides as well as their catalytic activity.

Various cases of metal doping have been reported to successfully improve the catalytic activities compared with the pristine metal oxide catalysts due to advantageous modification of electronic structure.<sup>157</sup> Interestingly, among these cases, the



**Fig. 9** (a) Cyclic voltammety of  $\text{Ni}_{1-x}\text{Fe}_x\text{O}_y\text{H}_z$  with different amount of Fe incorporation. The inset depicts a possible schematic of Fe incorporation into a  $\text{NiO}_x\text{H}_y$  platelet as a function of cycling in Fe saturated solution. Reproduced with permission.<sup>160</sup> Copyright 2017, American Chemical Society. (b) XAS-derived structural motifs prevalent during OER catalysis at high and intermediate Ni-content. (c) Simplified scheme of the electrochemical water splitting cycle with metal oxidation rate constant,  $k_{M,ox}$ , and the catalytic OER rate constant,  $k_{OER}$ . Reproduced with permission.<sup>164</sup> Copyright 2016, American Chemical Society. (d) d-Orbital configurations of iron and nickel cations at the surface. (e) The spin channel of Fe sites and Ni sites in  $\text{Ni}_x\text{Fe}_{1-x}\text{OOH}$  and  $\text{NiOOH}$ . (f) The orbital interactions between cations and the OER intermediates. Reproduced with permission.<sup>166</sup> Copyright 2020, Wiley-VCH.

incorporation of Fe can significantly enhance the OER catalytic activities of Co<sup>158</sup> or Ni (oxy)-hydroxides.<sup>73</sup> The catalytic activity of Ni<sub>x</sub>Fe<sub>1-x</sub>OOH even reached 1000-fold higher than the pristine NiOOH.<sup>20</sup> Nowadays, layer-structured Ni<sub>x</sub>Fe<sub>1-x</sub>OOH electrocatalysts show the best OER catalytic activity in alkaline condition. However, although the superior performance of Ni<sub>x</sub>Fe<sub>1-x</sub>OOH has been well accepted, there still remains a confusion how iron dopants influence the OER activity. In the 1980s, Corrigan<sup>71</sup> found the crucial effect of Fe impurities on the improvement of the catalytic activity, which has been later verified by Trotochaud and co-workers<sup>73</sup> in comparative experiments using an increasing Fe concentration in the KOH solution for the electrochemical tests. The group of authors argued that the incorporation of Fe promoted the electronic conductivity of NiOOH. As illustrated in Fig. 9a, the results of electrochemical investigation revealed that the Ni redox peaks of Ni<sub>x</sub>Fe<sub>1-x</sub>OOH catalysts shift toward the anodic direction with increasing *x*, suggesting the changes happened on the electronic structure.<sup>44,159,160</sup> Similar findings were also obtained from electrochemical testing of Fe-doped CoOOH;<sup>158</sup> it seemed that Fe incorporation had a similar influence on Ni and Co hosts. To further uncover the effect of Fe in Ni<sub>x</sub>Fe<sub>1-x</sub>OOH, Friebel and co-workers used operando X-ray absorption spectroscopy (XAS) to reveal that Fe<sup>3+</sup> in Ni<sub>x</sub>Fe<sub>1-x</sub>OOH occupies octahedral sites with unusually short Fe–O bond distances induced by edge-sharing with surrounding [NiO<sub>6</sub>] octahedra and they also identified that Fe cations were the active sites

instead of Ni cations by computational results.<sup>40</sup> In addition, Ahn and Bard utilized surface interrogation scanning electrochemical microscopy (SI-SECM) technology to directly investigate the OER kinetics of Ni<sup>4+</sup> and Fe<sup>4+</sup> in Ni(OH)<sub>2</sub>, FeOOH, and Ni<sub>x</sub>Fe<sub>1-x</sub>OOH and they claimed that the OER kinetics were much more fast on iron sites. These results seem to indicate that Fe dopants are the active sites.<sup>161</sup> In order to more directly identify the existence of Fe species in Ni<sub>x</sub>Fe<sub>1-x</sub>OOH, Mössbauer spectroscopy, has been applied in the investigation of Ni<sub>x</sub>Fe<sub>1-x</sub>OOH electrocatalysts.<sup>162</sup> For example, Jamie and co-workers reported the first direct evidence for the formation of Fe<sup>4+</sup> in NiFe oxide through operando Mössbauer spectroscopy.<sup>163</sup> However, they proposed that the detected Fe<sup>4+</sup> species did not directly serve as the active site in the OER process, although their existence had great effect on the improvement of OER activity for NiFe oxide electrocatalyst. Using *in situ* differential electrochemical mass spectrometry (DEMS) and XAS, Görlin *et al.*<sup>164</sup> found that Ni ions remained in the Ni<sup>2+</sup> state with the iron incorporation which might be responsible for the enhanced faradaic efficiency, while the Ni ions in pure NiOOH would be oxidized from Ni<sup>2+</sup> to Ni<sup>3+</sup> or Ni<sup>4+</sup> during OER process, leading to a lower faradaic efficiency (Fig. 9b and c). However, later, Nocera *et al.* proposed that the incorporation of Fe could promote the formation of Ni<sup>4+</sup> *via* the Lewis acid effect, causing the observed improved OER performance.<sup>165</sup> More recently, Xu and co-workers explained the effect of iron of Ni<sub>x</sub>Fe<sub>1-x</sub>OOH on the OER process from a perspective of the local spin

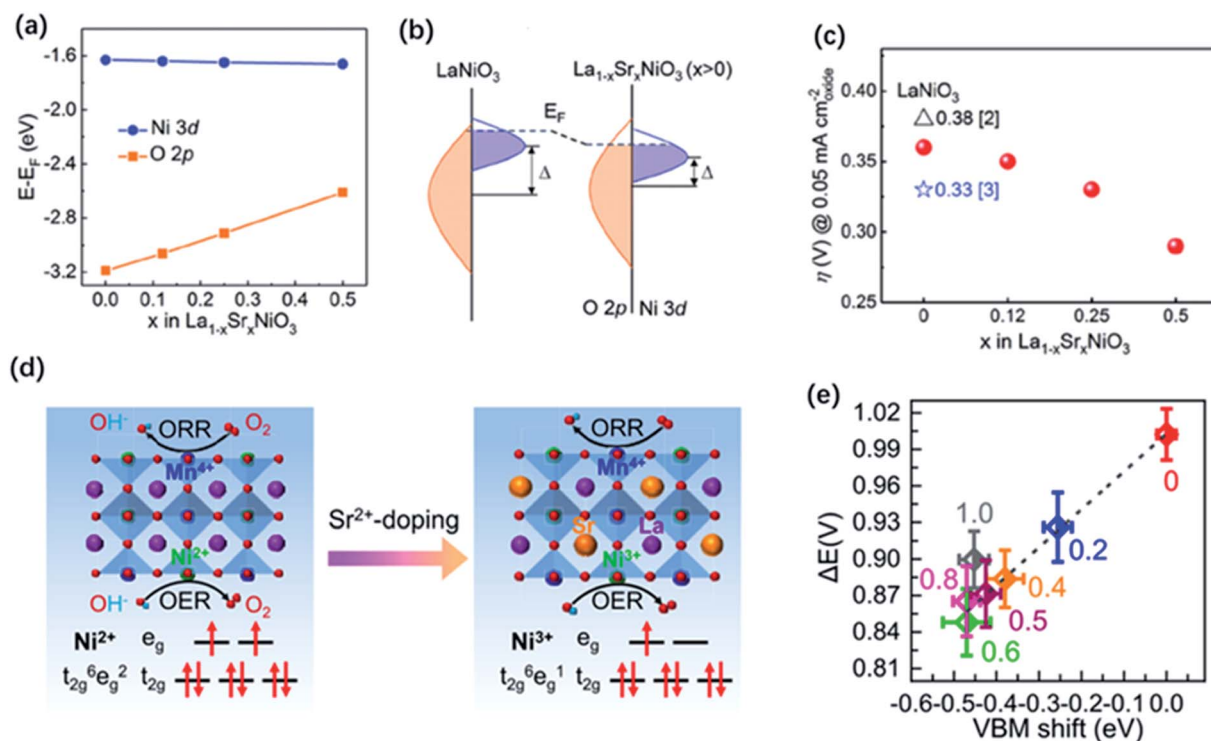


Fig. 10 (a) Average onsite energies of O 2p and Ni 3d orbitals with  $E_F$  as the reference. (b) Schematic energy band diagram for  $\text{LaNiO}_3$  and  $\text{La}_{1-x}\text{Sr}_x\text{NiO}_3$  ( $x > 0$ ), where  $\Delta$  is the charge transfer energy. (c) Overpotentials ( $\eta$ ) for LSNO films as required to obtain  $0.05 \text{ mA cm}^{-2}$  oxide in CV measurements. Reproduced with permission.<sup>173</sup> Copyright 2019, Wiley-VCH. (d) Double perovskite crystal structure of LNMO and LSNMO, and the corresponding electron configuration of Ni<sup>2+</sup> and Ni<sup>3+</sup> in the oxygen octahedral crystal field. (e)  $\Delta E$  values as a function of VBM shift. Reproduced with permission.<sup>98</sup> Copyright 2021, American Chemical Society.

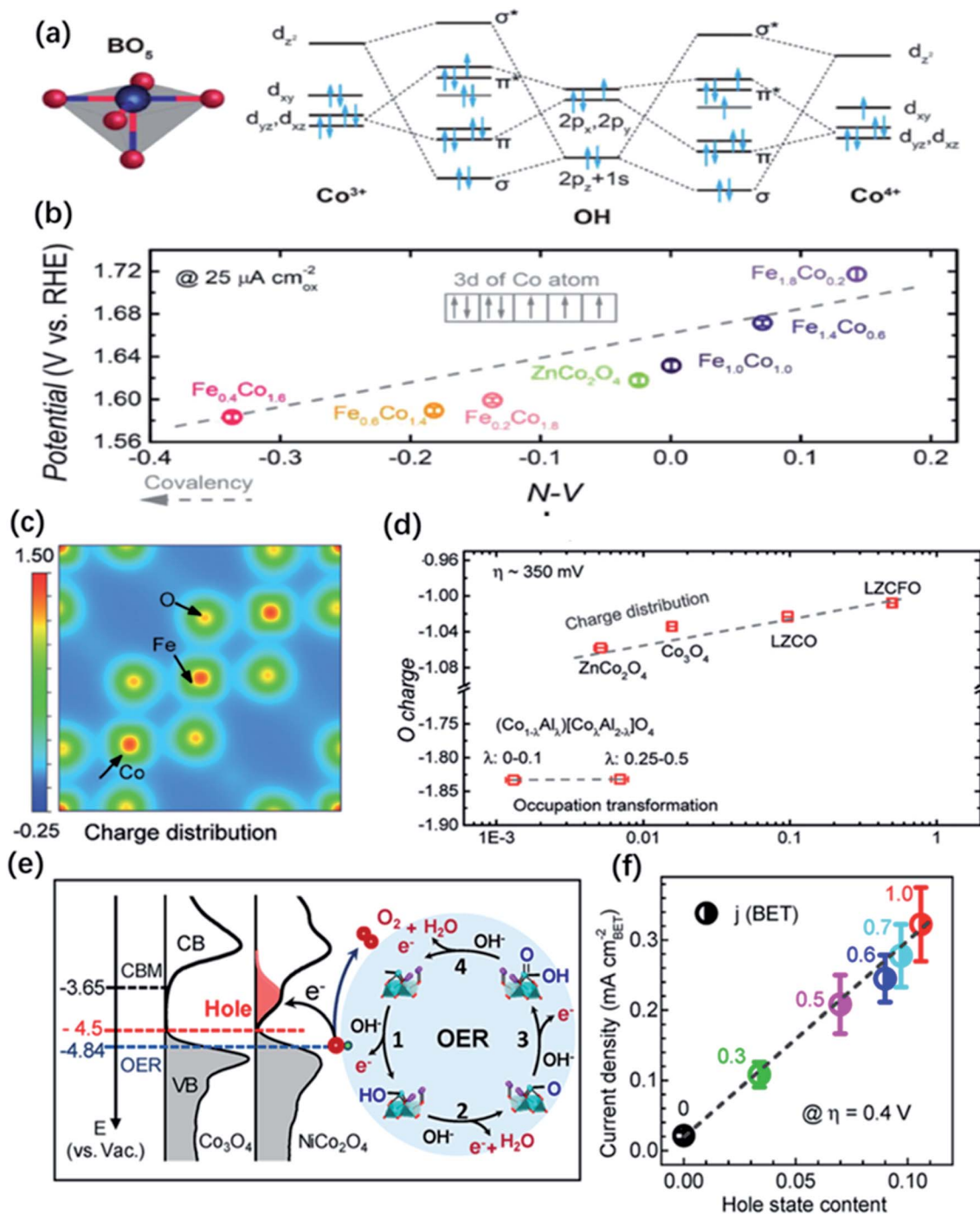


Fig. 11 (a) Molecular orbitals diagrams for the Co–OH (Co<sup>3+</sup>, Co<sup>4+</sup>) bonding at the surface of spinel oxides. (b) The OER activity, evaluated by the potentials at a current density of 25 μA cm<sup>-2</sup>, as a function of the N–V parameter. Reproduced with permission.<sup>129</sup> Copyright 2018, Wiley-VCH. (c) Illustration of charge density on [001] direction of LZFCO. (d) OER activity (specific current at an overpotential of ~350 mV) as a function of computed oxygen charges. Reproduced with permission.<sup>177</sup> Copyright 2019, Wiley-VCH. (e) Left panel: the measured occupied and unoccupied density of states (DOS) near the E<sub>F</sub> for Co<sub>3</sub>O<sub>4</sub> and NiCo<sub>2</sub>O<sub>4</sub>; the energy level is relative to vacuum level (vs. Vac.). Right panel: schematics of the four-step reaction pathway of OER. (f) Correlation of specific OER activity (current density at η = 0.4 V) with the amount of hole state extracted from O K edge XAS spectra. Reproduced with permission.<sup>68</sup> Copyright 2019, American Chemical Society.

configurations of surface iron and nickel.<sup>166</sup> As shown in Fig. 9d, the electronic configuration of surface Fe<sup>3+</sup> cations under the stable high spin state have 5 unpaired electrons (1 per d orbital), while the low spin state of surface Ni<sup>3+</sup> only possesses one single

electron at e<sub>g</sub> orbitals; a spin channel around the Fe sites can be obviously observed in Fig. 9e. During the OER process, the possible orbital interactions between Fe<sup>3+</sup>/Ni<sup>3+</sup> and the \*OH, \*O, and \*OOH species are shown in Fig. 9f. The group of

authors found that the different d-orbital occupations of  $\text{Fe}^{3+}$  and  $\text{Ni}^{3+}$  could result in distinct reaction kinetics, and that iron sites in  $\text{Ni}_x\text{Fe}_{1-x}\text{OOH}$  could promote the OER process through both the reactant adsorption and product desorption. Actually, other metal cations such as La, Y, Cd, Ce, Cr, and Zn *etc.*, also have been reported to be dopants incorporated in layered metal hydroxides, but their improvement on the OER performance has been less pronounced compared to Fe.<sup>167–170</sup>

A typical group of metal oxides featured by the highly compositional and structural flexibility with the unit formula of  $\text{ABO}_3$ , where A usually is the larger ions such as rare-earth metal cations and alkali-earth metal and B is the smaller ions such transition metal cations, is the perovskite-type metal oxides. The  $\text{ABO}_3$  structure of perovskite can greatly adapt cation incorporation through partially substituting the A or B cation or both of them with other metals giving the  $(\text{A}_x\text{A}'_{1-x})\text{BO}_3$ ,  $\text{A}(\text{B}_y\text{B}'_{1-y})\text{O}_3$ ,  $(\text{A}_x\text{A}'_{1-x})(\text{B}_y\text{B}'_{1-y})\text{O}_3$  or even more complex compositions. In the periodic table, over 90% of the metal elements could be potentially used to construct perovskite oxides, which provides a broad diversity and possibility for doping.<sup>171</sup> Depending on the adjustment of doping, perovskite oxides not only offer controllable electronic structure for modifying the OER activities but they also provide excellent models for establishing correlations between the activity and the electronic properties of surface and/or bulk properties *via* establishing electronic descriptors. For example, Xie and co-workers have reported that the Yb doped  $\text{CaMnO}_3$

electrocatalyst showed a significant enhancement of the OER performance compared with pure  $\text{CaMnO}_3$ , which could be ascribed to optimized electron occupation numbers of  $e_g$  orbital by doping.<sup>172</sup> Zhao and co-workers also found that the co-doping of Sr and Fe into  $\text{PrBaCo}_2\text{O}_{5+\sigma}$  perovskites led to a favourable  $e_g$  electron filling, which could improve the intrinsic OER activity.<sup>109</sup> Mefford and co-workers synthesized a series of  $\text{La}_{1-x}\text{Sr}_x\text{CoO}_{3-\delta}$  perovskites with the different amount of Sr dopants. They claimed that the covalency of the Co–O bond can be enhanced through Sr incorporation and thus influence the OER activities.<sup>36</sup> Similarly, Xu and co-workers introduced Fe dopants into  $\text{LaCoO}_3$  for modifying its OER activity,<sup>138</sup> and they showed by means of DFT calculation that the increased performance relates to an enlarged covalency of Co–O with 10% incorporation of Fe in  $\text{LaCoO}_3$ . On the other hand, Liu and co-workers analysed the effect of Sr doping in the  $\text{LaNiO}_3$  on the improvement of OER activity, and they rationalised the observed results based on the analysis of energy levels derived from X-ray photoemission spectroscopy (XPS) and X-ray absorption spectroscopy (XAS) in combination with DFT calculation.<sup>173</sup> As is shown in Fig. 10a–c, the results indicated that an up-shift energy of the O 2p band relative to the Fermi level occurred with rising  $x$  in  $\text{La}_{1-x}\text{Sr}_x\text{NiO}_3$ , which strengthened the degree of Ni 3d–O 2p hybridization and reduced the charge transfer energy, which resulted in a significant enhancement the OER activity. IN a similar analysis approach, Zhang and co-workers recently tuned the electronic structure of

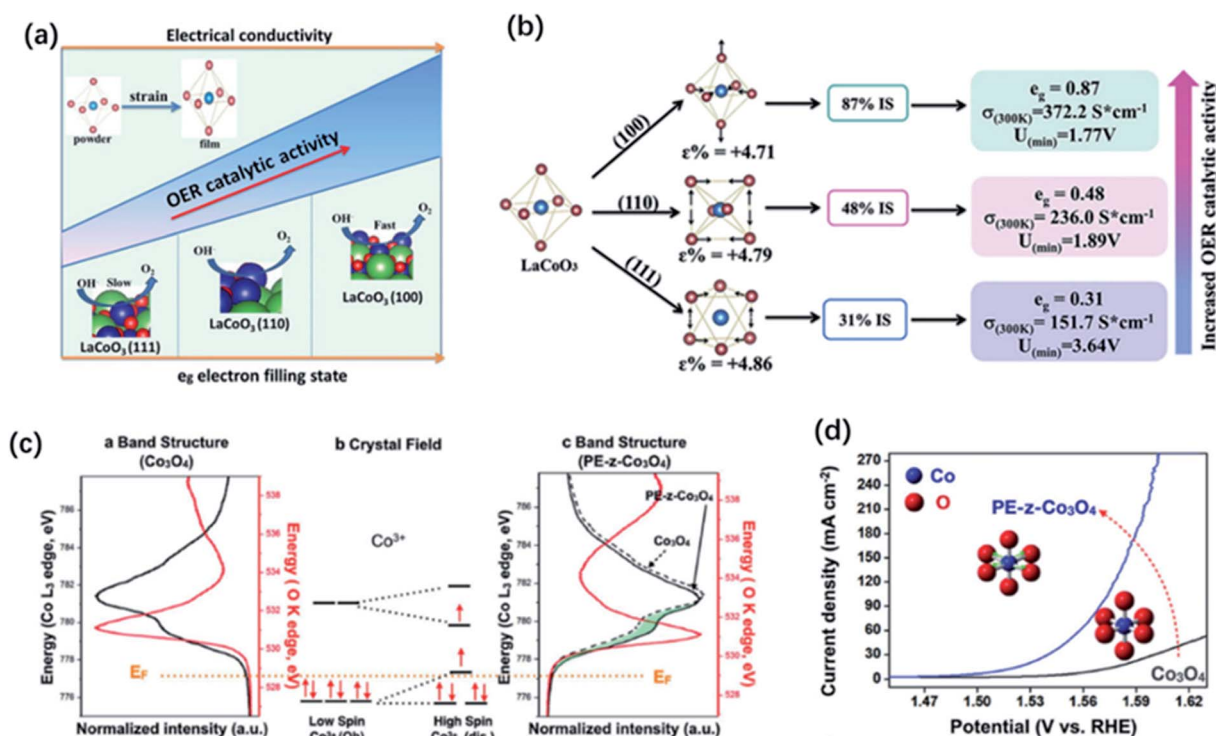


Fig. 12 (a) Spin-state regulation by lattice-oriented control of  $\text{LaCoO}_3$  epitaxial films. (b) The relationship between OER activity and spin configuration (free energy, conductivity as well as  $e_g$  electrons filling status) of differently oriented  $\text{LaCoO}_3$  films. Reproduced with permission.<sup>58</sup> Copyright 2017, Elsevier Inc. (c) Illustration of Co 3d and O 2p bonding orbitals based on Co (K edge) and O (L edge) spectra and (b) crystal field of cobalt ions with different spin states of  $\text{Co}^{3+}$ . (d) Linear sweep voltammetry (LSV) curves of  $\text{Co}_3\text{O}_4$  and  $\text{PE-z-Co}_3\text{O}_4$ . Reproduced with permission.<sup>181</sup> Copyright 2018, Wiley-VCH.



double perovskite  $\text{La}_2\text{NiMnO}_6$  by Sr doping, which effectively led to higher OER activities of  $\text{La}_{2-x}\text{Sr}_x\text{NiMnO}_6$ .<sup>98</sup> Their investigation on electronic structures discovered that the Sr doping induced  $\text{Ni}^{3+}$  states, resulting in the upshifted valence band centre and enhanced the hybridization of O 2p with Ni 3d and Mn 3d orbitals (Fig. 10d and e). Besides, the increased hole states originated from  $\text{Ni}^{3+}$  states decreased the energy barrier of the electron transfer from 0.44 to 0.12 eV. They claimed that the enhanced hybridization of O 2p with Ni and Mn 3d orbitals and the increased hole states should be responsible for the better OER activities.

Another group of metal oxides characterized by controllable composition and electron structure is the spinel family. In the typical structure of spinel oxides, there exist two geometric sites including tetrahedral (Td) sites which are normally half occupied by metal cations and octahedral (Oh) sites of which only one-eighth are filled by metal cations. The unoccupied geometric sites enable spinel-type oxides to largely accommodate the migration of metal ions. Otherwise, two geometric TM cations possess different coordination environment and they would compete for overlapping with the neighbouring oxygen because four neighbouring metal cations (three are in octahedral sites and the other one is in the tetrahedral sites) shared one common oxygen atom. Thus, the electronic properties of the metal cations in both sites and the oxygen anions closely correlate with each other. Doping or substituting other metal ions into Oh or Td sites becomes an effective approach to tuning the electronic structure of spinel oxides and influences their OER performance. For example, Zn, Ni, Cu, and Fe dopants introduced into  $\text{Co}_3\text{O}_4$  spinel have been widely reported and their effects on the electronic structure and catalytic properties have been deeply investigated.<sup>37,170,174,175</sup> Driess and co-workers found that the OER activity of  $\text{ZnCo}_2\text{O}_4$  electrocatalyst was slightly better than  $\text{Co}_3\text{O}_4$  in alkaline electrolyte, which could be ascribed to more accessible  $\text{Co}^{3+}$  in octahedral sites of  $\text{ZnCo}_2\text{O}_4$  surface owing to the preferential loss of Zn ions under reaction conditions.<sup>37</sup> Tan and co-workers synthesized different metal incorporated spinel catalysts  $\text{MCo}_2\text{O}_4$  ( $\text{M} = \text{Ni}, \text{Zn}, \text{Mn}, \text{etc.}$ ) and they uncovered that  $\text{ZnCo}_2\text{O}_4$  with a normal spinel structure where all the Co cations were in octahedral sites exhibited the better OER activity.<sup>176</sup> The degree of M–O covalency, as measured by the hybridization of M d and O 2p states, is also applied to connect the effect of doping with OER catalytic properties in spinel. Xu co-workers conducted an investigation on linking the Co–O covalency with the OER activity of Fe incorporated  $\text{ZnFe}_x\text{Co}_{2-x}\text{O}_4$  spinel catalyst.<sup>129</sup> The results showed that 10–30 at% Fe doping was beneficial to promote the injection/extraction of electrons from oxygen and lead to an enhanced covalency between Co 3d and O 2p, which is responsible for the improvement of the OER performance (Fig. 11a and b). Later, the same group reported that the OER activity followed the trend of  $\text{Co}_3\text{O}_4 > \text{ZnCo}_2\text{O}_4 > \text{AlCo}_2\text{O}_4$ , indicating that the octahedral  $\text{CoO}_6$  block dominated the OER activity, and explained the higher catalytic activities of the octahedral metal cations than the tetrahedral ones by the concept of covalency.<sup>177</sup> They also observed a greatly improved OER performance by the doping of Li and Fe into  $\text{ZnCo}_2\text{O}_4$ ,

obtaining a highly active  $\text{Li}_{0.5}\text{Zn}_{0.5}\text{Fe}_{0.125}\text{Co}_{1.875}\text{O}_4$  catalysts. They proposed that the enhanced OER performance was attributed to the charge shift from the oxygen to metal cations in octahedral sites *via* Li doping and enlarged covalency of  $\text{Co}_{\text{Oh}}\text{–O}$  bond *via* Fe doping (Fig. 11c and d). Additionally, Zhang and co-workers introduced Ni into  $\text{Co}_3\text{O}_4$  and conducted a systematic investigation on the influence of Ni dopants on the electronic structures of  $\text{Ni}_x\text{Co}_{3-x}\text{O}_4$ .<sup>68</sup> Using XPS and XAS, they found that the incorporation of Ni in the trivalent oxidation state induces the generation of an unoccupied hole state, which could decrease the energy barriers of electron transfer for a faster kinetics. Results also showed that there is a shift of the VBM by 0.27 eV closer to the  $E_F$  and enhanced the hybridization of O 2p with Ni 3d and Co 3d orbitals, which promoted the adsorption of OH intermediates on the surface of  $\text{Ni}_x\text{Co}_{3-x}\text{O}_4$  for a better OER activity than  $\text{Co}_3\text{O}_4$  (Fig. 11e and f).

In addition to the examples mentioned above, other studies on tuning the electronic structure and improving the OER activity of metal oxides by doping have also been reported. For instance, Zhang and co-workers modified the electronic properties of NiO through Li doping ( $\text{Li}_x\text{Ni}_{1-x}\text{O}$ ) to improve the performance toward OER.<sup>103</sup> According to their results, the significantly improved OER activities of  $\text{Li}_x\text{Ni}_{1-x}\text{O}$  could be ascribed to the generation of hole states at 1.1 eV above the Fermi level and the enlarged hybridization degree of O 2p–Ni 3d by Li doping which in turn resulted in the optimal adsorption strength of OH intermediates. Patzke *et al.* improved the OER activity of  $\alpha\text{-Mn}_2\text{O}_3$  through Mo doping.<sup>178</sup> According to their findings,  $\text{Mn}_2\text{O}_3$  with 2.64% Mo dopants exhibited the best OER activity. They claimed that the doping of Mo into the  $\alpha\text{-Mn}_2\text{O}_3$  lattice along with the structural distortion had great influence on improving the OER activity of pure  $\alpha\text{-Mn}_2\text{O}_3$ . Previous research showed that the energy of  $\text{Ni}^{\text{IV}}/\text{Ni}^{\text{III}}$  redox pair located at the top of the O-2p bands resulted in a strong bonding strength between the adsorbates on the surface and  $\text{Ni}^{\text{IV}}/\text{Ni}^{\text{III}}$  redox cation, which was an important premise for high OER activity. However, stoichiometric  $\text{LiNiO}_2$  with well-ordered  $\text{Li}^+$  and  $\text{Ni}^{\text{III}}$  is difficult to be obtained due to the generation of  $\text{Ni}^{\text{II}}$ . Based on these, Gupta and co-workers successfully prepared Al doped  $\text{LiNiO}_2$  catalysts ( $\text{Li}_{1-x}\text{Ni}_{1+x}\text{O}_2$ ) with better ordered lattice structure *via* solution combustion method at low temperature.<sup>179</sup> They proposed that Al dopants could stabilize  $\text{Ni}^{\text{III}}$  with higher concentration for the better OER activity.

## 5.2 Strain engineering

The compositional flexibility of metal oxides could provide convenience to modulate the electronic structure for a better OER activity, present advanced thin-film deposition techniques also offer the great opportunities to tune the properties of functional metal oxides through strain engineering. The physical and chemical properties of thin film and nanostructured metal oxides may largely differ from their bulk or powder form, which results in extra degrees of freedom for the modification by strain engineering. The modulation of electronic structure through strain effects can change a number of electronic structure parameters such as the metal d-band and oxygen 2p-

band centres and  $e_g$  occupancy *etc.* In the spinel and perovskite oxides, strain can lead to the distortion of metal–oxygen octahedra from the ideal cubic symmetry. The degeneracy of  $e_g$  orbitals can be changed through enlarging the electron filling in the  $d_{x^2-y^2}$  orbital caused by tensile strain and in the  $d_{z^2}$  orbital caused by compressive strain. Owing to the alteration in orbital overlap, the metal d-band and oxygen 2p-band centers as well as the binding strength of the adsorbed intermediates can be changed. For example, Petrie and co-workers observed the improvement of the OER activity on compressively strained  $\text{LaNiO}_3$  thin film in comparison with their unstrained state.<sup>180</sup> They proposed that compressive strain could cause orbital asymmetry and induce splitting of the  $e_g$  orbitals which result in the higher occupancy of the out-of-plane  $d_{z^2}$  orbitals. The extra occupancy of the  $d_{z^2}$  orbitals reduces the binding strength between the metal cation and adsorbate intermediates, resulting in better activity for OER and ORR.

Recently, much attention has focused on the strain engineering of  $\text{LaCoO}_3$  perovskite to modify its electronic structure, due to unique nature that the electron configuration of  $\text{Co}^{3+}$  that can thermally transit from low spin (LS:  $t_{2g}^6e_g^0$ ) to higher spin state with the  $e_g$  orbital configuration of  $e_g^{\sim 1.0}$  at room temperature. By decreasing the particle size of  $\text{LaCoO}_3$  to 80 nm, Zeng *et al.* observed the spin-state transition of Co ions from low-spin to high-spin states and the increase of  $e_g$  filling from unity to close to the favourable electron number of 1.2. Thus,  $\text{LaCoO}_3$  catalysts with the particle size of 80 nm displayed a better OER activity than that of other larger sized samples and the bulk.<sup>57</sup> Similarly, through strain engineering, Xie and co-workers prepared three kinds of epitaxially strained thin films of  $\text{LaCoO}_3$  perovskite with (100), (110), and (111) lattice orientation, respectively.<sup>58</sup> The  $\text{LaCoO}_3$  films with different lattice orientations exhibited different distortion degree of the  $\text{CoO}_6$  octahedron, causing a spin-state transition occurred on cobalt from a low spin state (LS  $t_{2g}^6e_g^0$ ) to an intermediate spin state (IS  $t_{2g}^5e_g^1$ ). As a result,  $\text{LaCoO}_3$  films with (100) orientation showed the best OER performance duo to its favourable  $e_g$  electron number of 0.87 and higher electrical conductivity (Fig. 12a and b). Modulating the spin state of cobalt by strain control was also reported by Hsu and co-workers.<sup>181</sup> They found that enlarging

the lattice constant on the  $\text{Co}_3\text{O}_4$  surface could change the spin state of surface  $\text{Co}^{3+}$  cations from  $t_{2g}^6$  to the high spin state of  $t_{2g}^4e_g^2$ , which remarkably enhanced the overlap of the  $e_g$  orbital of cobalt with the oxygen adsorbates, and thereby improving the OER performance (Fig. 12c and d).<sup>181</sup>

Strain could also enhance the formation of oxygen vacancy to modify the OER activity of metal oxide catalysts. For  $\text{SrCoO}_{3-\delta}$  thin films, tensile strain would upshift the O 2p-band relative to the Fermi level and it is beneficial for the formation of oxygen vacancy manifesting as the reduction of a small amount of  $\text{Co}^{4+}$  to  $\text{Co}^{3+}$ . Lee *et al.* reported that the oxygen deficiencies of  $\text{SrCoO}_{3-\delta}$  thin films can be tailored from  $\delta \leq 0.1$  to  $\delta \sim 0.25$  through epitaxial strain engineering.<sup>182</sup> The changes in the oxygen content happened on the modest amounts of tensile strain over 1%. The electrochemical measurement showed the improved OER activities of tensile-strained  $\text{SrCoO}_{3-\delta}$  thin films with an increased degree of strain. The authors ascribed the enhancement of OER activity to the existence of oxygen vacancies or the increase of  $\text{Co}^{3+}$  caused by strain. Otherwise, Du *et al.* also confirmed the effect of strain on oxygen vacancy formation.<sup>183</sup> They grew the tensile strained  $\text{NdNiO}_3$  thin films on  $\text{SrTiO}_3$  substrates and found that enhancing tensile strain would generate a higher content of oxygen vacancies in  $\text{NdNiO}_3$ . They claimed that when the effect of oxygen vacancies could be negligible, the tensile strain was not good for the OER process while compressive strain was beneficial for that due to the strain-induced orbital splitting for a higher filling of the  $d_{3z^2-r^2}$  orbital and resultant weaker Ni–O binding strength. Nevertheless, when the tensile strain was large enough to enhance the formation of oxygen vacancies, an improvement of the OER performance could be seen. The improved performance was associated with an  $e_g$  orbitals occupancy slightly larger than 1 due to the partial reduction of  $\text{Ni}^{3+}$  to  $\text{Ni}^{2+}$  caused by the oxygen vacancy formation.

### 5.3 Defect engineering

The structural and compositional flexibility of metal oxides also involve high tolerance in element defects, including metal cations and oxygen vacancies.<sup>184,185</sup> Defects can greatly influence the electron and orbital distribution of metal oxides. Thereby,

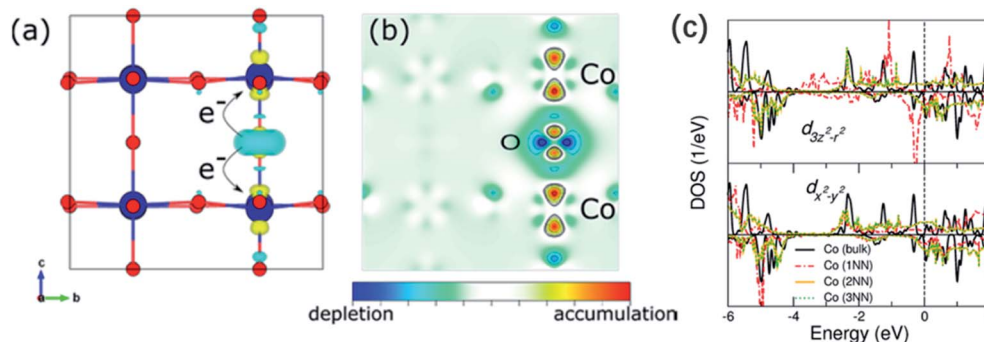


Fig. 13 (a) and (b) Differential charge density upon the creation of oxygen vacancies and a slice along the (100) plane. (c)  $d_{3z^2-r^2}$  and  $d_{x^2-y^2}$  contributions of the Co ions in a pristine bulk system and the Co ions surrounding oxygen vacancies. Reproduced with permission.<sup>189</sup> Copyright 2016, American Chemical Society.

defect engineering is an effective strategy to modify the electronic structure and catalytic properties of metal oxide electrocatalysts. For example, Wang *et al.* produced Sn vacancies on the surface of SnCoFe perovskite hydroxide ( $\text{SnCo}_{0.9}\text{Fe}_{0.1}(\text{OH})_6$ ) by energetic Ar plasma treatment.<sup>90</sup> The generated Sn vacancies reduced the coordination numbers of the neighbouring active sites, which promoted reactant adsorption and charge transfer. And the Sn vacancies also promoted exposure of active Co and Fe sites, tuned conductivity and adsorption energy of the intermediates, thus significantly improving the OER performance. Zhu and co-workers reported the introduction of A-site cation defects in  $\text{La}_{1-x}\text{FeO}_{3-\delta}$ , which lead to the formation of highly oxidative oxygen species ( $\text{O}_2^{2-}/\text{O}^-$ ) and  $\text{Fe}^{4+}$  species.<sup>156</sup> They proposed that the  $\text{O}_2^{2-}/\text{O}^-$  species might be the active sites for OER and  $\text{Fe}^{4+}$  species with an optimal  $e_g$  orbital filling ( $t_{2g}^3e_g^1$ ) also promoted the OER processes.

Actually, generating oxygen vacancies is easier compared to metal cation vacancies owing to the low formation energy of oxygen vacancies and they usually have a direct influence on the M–O bonds.<sup>186</sup> Therefore, many studies focused on the introduction of oxygen vacancies in metal oxides to tune their electronic structure and improve the OER performance. Wang *et al.* reported that the introduced oxygen vacancies in Zn-doped  $\text{CoOOH}$  would cause the transformation of the regular  $\text{CoO}_6$  octahedron to an elongated or compressed  $\text{CoO}_6$  octahedron (Jahn–Teller distortion), which enhance the lift degeneracy of the  $t_{2g}$  and  $e_g$  orbitals.<sup>185</sup> The electron configuration of  $\text{Co}^{3+}$  ( $t_{2g}^6e_g^0$ ) transferred to a more optimized  $e_g$  state filling ( $t_{2g}^5e_g^1$ ), which could facilitate bonding with intermediates due to the stronger overlap of  $\sigma$ -bonding  $e_g$  with O-2p than  $\pi$ -bonding  $t_{2g}$  orbital and finally promote the OER activity. Kim and co-workers prepared the oxygen deficient perovskite  $\text{Ca}_2\text{Mn}_2\text{O}_5$  with a higher OER mass activity of  $30.1 \text{ A g}^{-1}$  at  $1.70 \text{ V}$  (vs. RHE) in comparison to the perovskite  $\text{CaMnO}_3$ .<sup>187</sup> They claimed that the unit cells of  $\text{Ca}_2\text{Mn}_2\text{O}_5$  were elongated in *b*- and *c*-directions owing to the existence of oxygen vacancies and the electronic configuration of  $\text{Mn}^{4+}$  ( $3d^3$ ) with  $t_{2g}^3$  was optimized to  $t_{2g}^3e_g^1$  of  $\text{Mn}^{3+}$  ( $3d^4$ ), which both were favorable for the adsorption of OH intermediates, thereby exhibiting higher OER activity. Moreover, Mefford *et al.* found that the increase of Sr substitution in  $\text{La}_{1-x}\text{Sr}_x\text{CoO}_{3-\delta}$  resulted in more oxygen vacancies, stronger covalency of the Co–O bond, which led to better OER performance.<sup>36</sup>

Zhu and co-workers found that controlling oxygen defects in  $\text{PrBa}_{0.5}\text{Sr}_{0.5}\text{Co}_{1.5}\text{Fe}_{0.5}\text{O}_{5+\delta}$  (PBSCF) could obviously promote the OER kinetics.<sup>188</sup> Using surface-sensitive characterization and DFT calculations, they proved that excessive oxygen vacancies enhanced  $\text{OH}^-$  affinity and lowered the theoretical formation energy of  $\text{O}^*$  intermediates on the surface, thereby significantly improving the OER activity. But they also proposed that excessive oxygen vacancies may lead to larger energy band gap and lower O 2p band position relative to the Fermi level of  $\text{PrBa}_{0.5}\text{Sr}_{0.5}\text{Co}_{1.5}\text{Fe}_{0.5}\text{O}_{5+\delta}$ , which may be not beneficial for OER process. Through DFT calculations, Tahini and co-workers found that the introduction of oxygen vacancies in  $\text{SrCoO}_3$  would lead to a local distortion of the Co–O bonds and donate its electrons to neighbouring Co cations.<sup>189</sup> The reduced Co sites

on the surface and Co–O bonds in the bulk lead to a reduction in the overlap between the Co 3d and O 2p orbitals, thereby resulting in the decrease of the electron transfer between the catalyst and the reactant intermediates (Fig. 13a–c). In addition, Miao *et al.*<sup>190</sup> and Liu *et al.*<sup>191</sup> also found that OER activities of  $\text{PrBaCo}_2\text{O}_{6-\delta}$  and  $\text{La}_{0.7}\text{Sr}_{0.3}\text{CoO}_{3-\delta}$ , respectively, decreased after introducing oxygen vacancies. Thus, optimizing OER performance by properly introducing oxygen vacancies is very important.

## 6. Summary and perspectives

Improving the kinetics of the OER with low-cost TM oxide catalysts is one of the big challenges of the scientific and engineering communities owing to the important role of the electrochemical water splitting in the introduction of renewable sources in any future energy scenario. In this article, we reviewed studies of the electronic structure of the TM oxide catalysts for the OER and the design of materials with high catalytic performances on basis of such electronic properties. The identification of electronic structure parameters associated with (i) the adsorbate binding strength, (ii) the covalency of the TM–O bonds, and (iii) the interface energetics can provide rational bases to explain the influence of the composition on the OER catalytic activity and guidelines for the formulation of new catalyst with improved activity. A key observation in the current literature is that highly covalent oxides tend to have higher catalytic activities, triggering a mechanism in which the lattice oxygen has been shown to participate. This has brought the spotlight on the activation of lattice oxygen as a way to improve the OER catalytic activity of TM oxides.<sup>127,130,148,151,152</sup>

Elucidating the plausible involvement of lattice oxygen in highly active oxide catalysts requires the development of novel research approaches to address two important scientific issues: (1) the complex reaction mechanisms and (2) the stability of the TM oxide catalysts. This has motivated the development of *in situ* and operando characterisation technique such as isotope labelling combined with differential electrochemical mass spectrometry, *in situ* IR and Raman spectroscopies aiming to capture chemical information of reaction intermediates in mechanistic studies.<sup>147</sup> The study of the chemical state and the structure of the catalysts during the OER process has also been studied with operando XAS.<sup>192</sup> Additional *in situ* characterisation studies with techniques such as ambient pressure X-ray photoelectron spectroscopy (NAP-XPS) to study the chemistry and energetics of the solid–liquid interface,<sup>193,194</sup> may bring new opportunities to obtain a more realistic understanding of the relationship between electronic structure, adsorbate binding strength, and the OER catalytic activity.

The involvement of lattice oxygen in the OER of highly active catalyst leads to surface reconstructions and may also decrease the stability of the catalysts, especially at higher overpotentials. Therefore, the stability of the catalyst remains an important scientific issue in this field. Studying surface reconstruction phenomena could benefit from the use of well define surfaces of model samples. This experimental approach has been recently used in the description of mechanisms that connect the lattice

oxygen activation, metal dissolution, and amorphization in TM OER electrocatalysts.<sup>195</sup> Compared with ill-defined porous electrodes commonly used in OER studies, single crystal thin films permit a high degree of control over crystal facets, defects, and compositions, which could be used as model catalysts to gain a definitive structure–property relationships at a molecular level, and the study of surface reconstructions in a controlled way.

## Conflicts of interest

There are no conflicts to declare.

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