


Cite this: *RSC Adv.*, 2021, 11, 14063

Catalytic activity of Co-nanocrystal-doped tungsten carbide arising from an internal magnetic field†

M. Morishita,^a A. Nozaki,^a H. Yamamoto,^a N. Fukumuro,^a M. Mori,^b K. Araki,^b F. Sakamoto,^b A. Nakamura^b and H. Yanagita^c

Pt is an excellent and widely used hydrogen evolution reaction (HER) catalyst. However, it is a rare and expensive metal, and alternative catalysts are being sought to facilitate the hydrogen economy. As tungsten carbide (WC) has a Pt-like occupied density of states, it is expected to exhibit catalytic activity. However, unlike Pt, excellent catalytic activity has not yet been observed for mono WC. One of the intrinsic differences between WC and Pt is in their magnetic properties; WC is non-magnetic, whereas Pt exhibits high magnetic susceptibility. In this study, the WC lattice was doped with ferromagnetic Co nanocrystals to introduce an ordered-spin atomic configuration. The catalytic activity of the Co-doped WC was ~30% higher than that of Pt nanoparticles for the HER during the hydrolysis of ammonia borane (NH_3BH_3), which is currently attracting attention as a hydrogen fuel source. Measurements of the magnetisation, enthalpy of adsorption, and activation energy indicated that the synergistic effect of the WC matrix promoting hydrolytic cleavage of NH_3BH_3 and the ferromagnetic Co crystals interacting with the nucleus spin of the protons was responsible for the enhanced catalytic activity. This study presents a new catalyst design strategy based on the concept of an internal magnetic field. The WC–Co material presented here is expected to have a wide range of applications as an HER catalyst.

Received 12th February 2021
Accepted 1st April 2021

DOI: 10.1039/d1ra01181b
rsc.li/rsc-advances

Introduction

The development of catalysts for the hydrogen evolution reaction (HER) is critical for producing hydrogen fuel as a substitute to fossil fuels to reduce global CO_2 emissions. Tungsten carbide (WC) has 10 valence electrons, $5d^6$ from W and $2p^4$ from C, similar to Pt ($5d^{10}$). Since 1973, when Levy and Boudart¹ proposed that WC exhibited singular catalytic activity similar to Pt, this topic has been investigated by many researchers. Bennett *et al.*² measured the valence band spectrum of WC using X-ray photoelectron spectroscopy (XPS) and compared it with the spectra of W and Pt. Near the Fermi level (E_F), the electronic density of states (DOS) of WC more closely resembled that of Pt³ than W.⁴ Houston⁵ compared the DOS of W, WC, and Pt using soft X-ray appearance potential spectroscopy (SXAPA). Contrary to the results of previous studies,^{3,4} the width of the unoccupied portion of the 5d band of W in WC was larger than in pure W.⁵ Another study⁶ showed that the differences between the DOS

measured by SXAPA⁵ and predicted using a rigid band model^{1,2} resulted in different crystal structures of body centred cubic (bcc) W, face centred cubic (fcc) Pt, and hexagonal WC. Mattheiss and Hamann⁷ investigated the DOS of the bulk and (0001) surface of WC using a relativistic linear augmented-plane wave method.⁸ The DOS near the E_F of the (0001) surface associated with the catalytic properties was larger than in the bulk. Recently, the electronic states of Pt were updated⁹ using the projector-augmented-plane wave method implemented in the VASP code⁸ using the generalised gradient approximation (Perdew–Burke–Ernzerhof version) (GGA-PBE),¹⁰ which clarified that E_F is located at the top of the 5d DOS, consistent with its highest catalytic activity of all elements.

Esposito *et al.*¹¹ showed that a monolayer film of Pt deposited on a bulk WC substrate exhibited a similar electrochemical hydrogen overpotential, η_{H_2} , to that of Pt in an aqueous H_2SO_4 solution, indicating that the surface electronic and chemical properties of monolayer Pt on a bulk WC substrate are significantly similar to those of bulk Pt. Zhang *et al.*¹² synthesised a $\text{Fe}_3\text{C}/\text{Co}_3\text{C}/\text{WC}/\text{C}$ carbide composite prepared by combining hydrothermal synthesis with resin impregnation and pyrolysis. Linear-sweep voltammetry of the $\text{Fe}_3\text{C}/\text{Co}_3\text{C}/\text{WC}/\text{C}$ in an aqueous KOH solution exhibited similar behaviour to that of a Pt/C electrode, indicating that it is a candidate cathode material for polymer electrolyte membrane fuel cells (PEMFCs).¹³ In addition, a 7.5 wt% Pt/ W_2C catalyst showed

^aDepartment of Chemical Engineering and Materials Science, University of Hyogo, 2167 Shosha, Himeji 671-2280, Japan. E-mail: morishita@eng.u-hyogo.ac.jp

^bGraduate Student of University of Hyogo, 2167 Shosha, Himeji 671-2280, Japan

^cSanalloy Industry Co., Ltd, 290-44 Takahashi, Fukusaki-cho, Kanzaki 679-2216, Japan

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1ra01181b



a current density 2–3 times higher than that of a commercial 20 wt% Pt/C for electrolysis in H_2SO_4 aqueous solution due to the synergistic effect of Pt and WC.¹³ Good catalytic activity for the hydrogen evolution reaction (HER) *via* electrolysis of the H_2SO_4 aqueous solution was observed using a powder micro-electrode composed of 16.20 wt% Pt/WC¹⁴ or a bimetallic carbide composed of Mo_2C and WC.¹⁵ Zheng *et al.* demonstrated that PEMFCs using a Nafion membrane containing WC nanoparticles exhibited enhanced power density and durability over 100 h of use.¹⁶

The catalytic activities of WC and its composites have only been observed under conditions of an applied external voltage,^{11–16} and the predicted intrinsic catalytic activity of mono WC under voltage-free operation has not yet been verified. Although Levy and Boudart observed that WC catalysed the reduction of WO_3 with hydrogen gas in the presence of water, and the isomerisation of 2,2-dimethylpropane to 2-methylbutane, the rates were merely 0.37% and 0.01%, respectively, in comparison with the performance of Pt.¹

Bennett *et al.*² noted that one of the intrinsic differences between WC and Pt is their magnetic properties; WC is non-magnetic, whereas Pt exhibits high magnetic susceptibility. Cerri *et al.*¹⁷ studied the magnetic properties of polycrystalline Gd and found that hydrogen chemisorption induced a dis-ordering of the electronic spin polarisation on the surface. An FTIR spectroscopy study confirmed that antiferromagnetic LaFeO_3 accelerated the nucleus spin conversion of ortho liquid hydrogen (composed of two antiparallel nucleus spins) to para liquid hydrogen (composed of two parallel nucleus spins) to reach equilibrium.¹⁸ Furthermore, Galces-Pineda *et al.*¹⁹ observed that an external magnetic field accelerated oxygen evolution during the electrolysis of water.

An application of HER catalysts is the generation of hydrogen fuel from NH_3BH_3 . In its stable crystal form, NH_3BH_3 contains 19.6 wt% hydrogen,²⁰ and is being investigated for efficient transportation of hydrogen-based fuel for portable fuel-cell systems. Previous studies investigated the HER by hydrolysis over 10 wt% Co²¹ or 2 wt% Pt²² (both supported by Al_2O_3) and found that the HER in the latter was significantly faster than that in the former. A similar HER in $\text{NH}_3\text{BH}_3(\text{aq})$ catalysed by Ni nanoparticles (NPs) supported by a zeolite molecular sieve was observed.²³

In this study, the WC lattice was doped with ferromagnetic Co nanocrystals ($\text{WC-Co}_{\text{carbide}}$)²⁴ to introduce an ordered-spin configuration as an internal magnetic field to avoid the need for an external applied voltage. The catalytic performance of $\text{WC-Co}_{\text{carbide}}$ for the HER in $\text{NH}_3\text{BH}_3(\text{aq})$ was compared with that of Pt nanoparticles (Pt_NPs), commercial mono WC, and a bcc W-Co solid solution ($\text{W-Co}_{\text{alloy}}$) in the spin glass state. The present study aims to investigate the (1) effect of an internal magnetic field instead of an external voltage on the catalytic activity; (2) kinetics of the HER; and (3) practical applications of the catalyst.

Materials and methods

The $\text{W-Co}_{\text{alloy}}$ powder consisting of a solid solution of Co-supersaturated bcc W was prepared using a hydrothermal

synthesis method that we have described previously.²⁴ The precursor materials were 99% ammonium tungstate pentahydrate ($5(\text{NH}_4)_2\text{O} \cdot 12\text{WO}_3 \cdot 5\text{H}_2\text{O}$; Kanto Chemical Co. Inc., Japan) and 99% cobalt acetate tetrahydrate ($\text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 4\text{H}_2\text{O}$; Kojundo Chemical Laboratory Co. Ltd, Japan), which were mixed to achieve a molar ratio of W : Co = 80 : 20. Briefly, $\text{W-Co}_{\text{alloy}}$ was carburised at 1173 K with a gas with a composition of 23 vol% CO_2 , 32 vol% H_2 , and 45 vol% Ar to form WC containing Co nanocrystals ($\text{WC-Co}_{\text{carbide}}$). In this study, CO_2 was used as the carburisation gas instead of CO to increase the safety of the process. Since CO_2 barely carburises the $\text{W-Co}_{\text{alloy}}$ based on the thermochemical equation, CO_2 was converted into CO *in situ* in the furnace using 50 : 50 mol% Fe–Al powder, where the Al thermally reduces CO_2 into CO. Although 10.8 ks was shown to be a sufficient processing time for carburisation using CO,²⁴ a total of 32.4 ks was required when CO_2 was used. During this carburisation procedure, the samples were ground at each 10.8 ks interval.

The chemical composition and structure of $\text{WC-Co}_{\text{carbide}}$ were confirmed by X-ray diffraction (XRD; Rigaku, Ultima IV) and electron probe microanalysis (EPMA; JOEL, JXA-8900) operated at a 15 kV accelerating voltage. The nanostructure of the Co crystal was observed by high-resolution transmission microscopy (HRTEM; JOEL, JEM-2100) operated at an accelerating voltage of 200 kV. The metallic state of the surface of the Co crystal was confirmed *via* XPS (ULVAC-PHI Inc., PHI5000) using monochromatic X-rays (Al $K\alpha$, 1486.6 eV). The specific surface area (SSA) of all samples was measured using the Brunauer–Emmett–Teller method (Shimadzu, TriStar II 3020) with the Kr physisorption isotherm obtained at 77 K.

A standard catalyst containing 1 wt% Pt_NPs on a carrier of Al_2O_3 particles (Catalysis Society of Japan)²⁵ was used to obtain comparative data for the catalytic reaction. The SSA of the Pt was $3.40 \text{ m}^2 \text{ g}_{\text{cat}}^{-1}$ and the total SSA of the catalyst, including the Al_2O_3 , was $176 \text{ m}^2 \text{ g}_{\text{cat}}^{-1}$.²⁵ A commercial mono WC particle (Kojundo Chemical Laboratory Co. Ltd, Japan; 99%) was also used as a reference sample.

The $\text{WC-Co}_{\text{carbide}}$, $\text{W-Co}_{\text{alloy}}$, Pt_NPs, and WC samples were flushed in a H_2 atmosphere at 473 K for 2 h. Then, 20 mg of each sample was placed in a glass test tube with 1 ml of H_2O . In addition, 0.5 mmol of $\text{NH}_3\text{BH}_3(\text{cr})$ was dissolved in 1.5 ml of H_2O to form an aqueous solution of $\text{NH}_3\text{BH}_3(\text{aq})$ that was mixed with each of the aforementioned samples to evolve hydrogen by hydrolysis of the solution. Four repeat measurements of the hydrogen evolution volume (HEV) were performed for the $\text{WC-Co}_{\text{carbide}}$, $\text{W-Co}_{\text{alloy}}$, and Pt_NPs samples, whereas only one measurement was considered sufficient for WC as it showed no catalytic activity.

The HER rate for each sample was evaluated by determining the slope of the HEV curve as a function of time using the least squares method. The average HER rates of the four repeat measurements are presented here. The 1-sigma error (68% confidence interval)²⁶ was calculated by dividing the standard deviation (σ) of the four measurements by the average value.

To determine E_a , the HEV over $\text{WC-Co}_{\text{carbide}}$ was measured at 308, 318, 328, and 338 K using a solution prepared with 1.0 mmol $\text{NH}_3\text{BH}_3(\text{cr})$ and 1.5 ml of H_2O . A higher solution



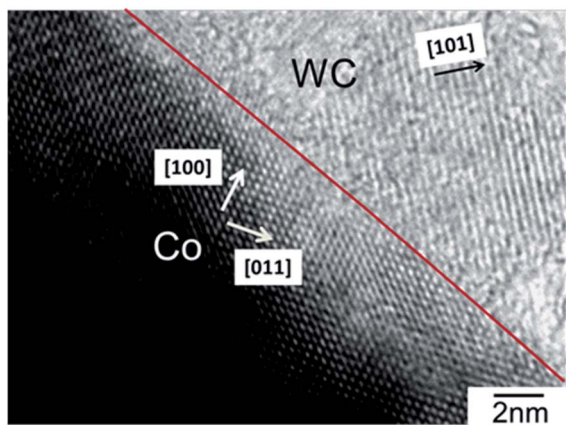


Fig. 1 Bright-field TEM image of WC-Co_{carbide}.

concentration than that used in the previous measurements was selected to achieve a longer reaction time and hence, more accurate Arrhenius plots. Three HER measurements were performed at each temperature and the average HER rate is presented. The 1-sigma error was evaluated by dividing 1σ of the three measurements by the average value at each temperature. The 1-sigma error was evaluated from the standard deviation of the 12 total measurements.

To determine the value of M of WC-Co_{carbide}, the magnetic hysteresis loop was measured using a SQUID instrument (QD, MPMS) under a magnetic field increased to a maximum of 3×10^4 Oe. Temperatures of 308 and 4 K were used to investigate the maximum value of M ,^{27,28} respectively.

The standard enthalpies of adsorption of $H_2(g)$, $\Delta_{ad}H_m^\circ$, for WC-Co_{carbide}, WC, and a commercial 99.9% Pt powder (Kojundo Chemical Laboratory Co. Ltd, Saitama) were measured at 423 K using a Calvet-type microcalorimetre (SETRAM, C80) in a H_2 atmosphere.

Results and discussion

Hydrogen evolution over WC-Co

Fig. 1 shows a representative bright-field TEM image of a WC-Co_{carbide} sample. Co crystals with the fcc structure, [100] orientation, and diameter of 60 nm were observed in the WC matrix with [101] orientation. The atomic configuration was the same as that observed in our previous study.²⁴ The SSA values of

the WC-Co_{carbide}, W-Co_{alloy}, and Pt_{NPs} samples are shown in Table 1. The WC-Co_{carbide} exhibited an SSA $\sim 60\%$ smaller than that of the Pt_{NPs}, whereas WC-Co_{carbide} exhibited a smaller SSA than that of W-Co_{alloy}.

Fig. S1 (see ESI†) shows the XRD results. Only peaks related to the WC matrix and Co crystals were observed. Fig. S2 (see ESI†) shows the Co 2p XPS spectrum of WC-Co_{carbide}. The peak at 778.4 eV was assigned to the metallic state Co^0 . Fig. S3 (see ESI†) shows representative EPMA/SEM images of W $L\alpha$, C $K\alpha$, and Co $K\alpha$ for the WC-Co_{carbide} particle. Because W, C, and Co were distributed homogeneously, HRTEM observation (Fig. 1) was necessary to distinguish the Co crystals from the WC matrix.

Fig. 2 shows the change in the average HEV over time (t) for hydrolysis of the $NH_3BH_3(aq)$ over WC-Co_{carbide}, W-Co_{alloy}, Pt_{NPs}, and commercial WC samples. The slope of the HEV(t) curve for WC-Co_{carbide} was steeper than that for Pt_{NPs}, indicating that WC-Co_{carbide} exhibited singular catalytic activity, whereas W-Co_{alloy} exhibited less activity.

The normalised HER rates per unit area, N_{HER} (H_2 mol $min^{-1} m^{-2}$), determined by normalising the HEV by the SSA values are shown in Table 1, along with their error values. The N_{HER} of WC-Co_{carbide} was $\sim 30\%$ higher than that of Pt_{NPs}, indicating that it exhibits singular catalytic activity similar to Pt. The N_{HER} of W-Co_{alloy} was only approximately 10% of that of WC-Co_{carbide}.

Activation energy

To clarify the mechanism of the singular catalytic activity of WC-Co_{carbide}, its E_a value was investigated. Fig. 3 shows the HEV as a function of time during the hydrolysis of $NH_3BH_3(aq)$ over WC-Co_{carbide} at 308, 318, 328, and 338 K. Comparing the data measured at 308 K to the corresponding curve in Fig. 2, the HEV at the same time was consistent with doubling the NH_3BH_3 concentration. Hence, a large number of reaction sites were available for studying the HER. The slopes of the HEV(t) curves increased with increasing temperature, indicating that the hydrolysis was a result of a thermally activated process.

The E_a is defined by the Arrhenius equation:

$$k_r = Ae^{-\frac{E_a}{RT}} \quad (1)$$

where k_r is the reaction rate constant, R is the gas constant ($8.3145 J K^{-1} mol^{-1}$), and A is the frequency factor indicating

Table 1 Summary of the SSA and HER results extracted from Fig. 2 and 3 for all sample types

Sample	SSA ($m^2 g_{cat}^{-1}$)	N_{HER} (H_2 mol $min^{-1} m^{-2}$)	N ($NH_3BH_3(aq)$ mmol $^{-1}$)	V (H_2O mL $^{-1}$)	T (K)
WC-Co _{carbide}	1.35	3.76 ± 0.37	0.5	2.5	308
WC-Co _{carbide}	1.35	7.63 ± 0.32	1.0	2.5	308
WC-Co _{carbide}	1.35	12.14 ± 2.42	1.0	2.5	318
WC-Co _{carbide}	1.35	26.92 ± 3.13	1.0	2.5	328
WC-Co _{carbide}	1.35	46.91 ± 7.37	1.0	2.5	338
W-Co _{alloy}	2.23	0.45 ± 0.11	0.5	2.5	308
Pt _{NPs}	3.40	2.90 ± 1.04	0.5	2.5	308
Comm. WC	1.03	0	0.5	2.5	308
Comm. Pt	0.12	—	—	—	—



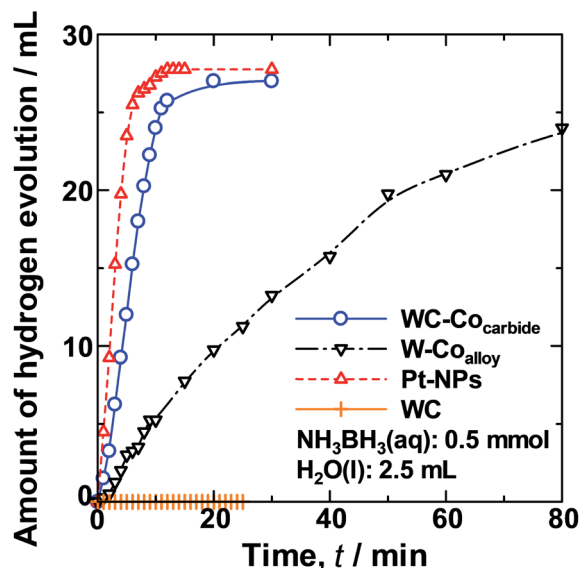


Fig. 2 Comparison of the HEV over time from hydrolysis of a $\text{NH}_3\text{BH}_3(\text{aq})$ solution (0.5 mmol, 2.5 ml) at 308 K using the various catalyst samples.

the number of collisions of the reaction species. The N_{HER} values summarised in Table 1 were regarded as k_r which were plotted as an Arrhenius plot in Fig. 4 ($\ln k_r$ vs. T^{-1}). The $\ln k_r$ values decreased linearly as a function of T^{-1} , where E_a was calculated from the slope of this curve. The E_a and A values for the HER over WC- $\text{Co}_{\text{carbide}}$ were $54.0 \pm 13.0 \text{ kJ mol}^{-1}$ and $1.01 \times 10^{10} \pm 2.44 \times 10^9 \text{ mol m}^{-2} \text{ min}^{-1}$, respectively, where the uncertainty was evaluated as the error propagation²⁶ of the uncertainties of the N_{HER} data at each temperature listed in Table 1. This E_a value was larger than the literature value for Pt NPs (2 wt% Pt on an Al_2O_3 support)²² and similar to that of Co

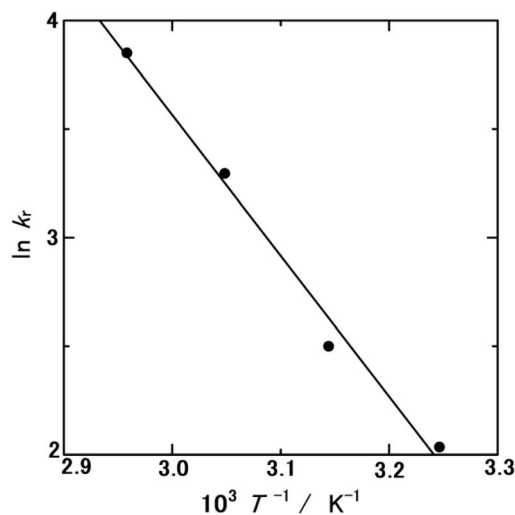


Fig. 4 $\ln k_r$ versus T^{-1} plot calculated from Table S1.†

NPs (10 wt% Co on an Al_2O_3 support),²¹ indicating that the hydrogen was released from the surface of the Co nanocrystals in the WC matrix.

The catalytic activity of WC- $\text{Co}_{\text{carbide}}$ was evaluated considering the hydrogen overpotential (η_{H_2}). The η_{H_2} of Pt is 0 V by definition, whereas for Co it is in the range of -0.25 to -0.47 V in 1 M H_2SO_4 aqueous solution at 303 K under a current density of $0.3\text{--}10 \text{ mA cm}^{-2}$.²⁹ Chandra and Xu²² determined E_a (21 kJ mol^{-1}) of the HER in $\text{NH}_3\text{BH}_3(\text{aq})$ over Pt NPs.²² The difference in the activation energy, ΔE_a , between WC- $\text{Co}_{\text{carbide}}$ and Pt NPs²² was 33 kJ mol^{-1} , which corresponds to -0.34 V, consistent with the η_{H_2} of Co.²⁹ This analysis indicated that the hydrogen release sites were the Co nanocrystals in the WC matrix. In general, η_{H_2} is defined as the complex energies consisting of the elementary processes: (I) alignment of $\text{H}^+(\text{aq})$ on the electrode; (II) accepting electrons to form the radical hydrogen atom $\text{H}(\text{rad.})$; (III) bonding two $\text{H}(\text{rad.})$ to form a H_2 molecule; (IV) convergence of the radicals to form hydrogen gas $\text{H}_2(\text{g})$; and (V) desorption of $\text{H}_2(\text{g})$ from the electrode. In this study, the separation of these elementary processes of η_{H_2} was impossible as E_a was determined only from a simple Arrhenius plot. Further computational simulations are necessary to clarify this point. The A value is discussed later with respect to the thermodynamic cycle.

Magnetic behaviour

The HER over the W- Co_{alloy} powder was not enhanced as observed for WC- $\text{Co}_{\text{carbide}}$ (Fig. 2), indicating that $\text{H}_2(\text{g})$ is not readily evolved over the Co atoms substituted into the bcc W lattice. However, the HER was enhanced in the case of the WC- $\text{Co}_{\text{carbide}}$ sample with Co atoms configured in the nanocrystal domain. The most likely difference between these configurations is related to their magnetic properties. Hence, M was investigated to clarify the different HER results. Fig. 5 shows M as a function of the magnetic field (H) for WC- $\text{Co}_{\text{carbide}}$ and W- Co_{alloy} at 4 K. WC- $\text{Co}_{\text{carbide}}$ showed a clear hysteresis loop, indicating that it is a ferromagnetic material. Assuming that the

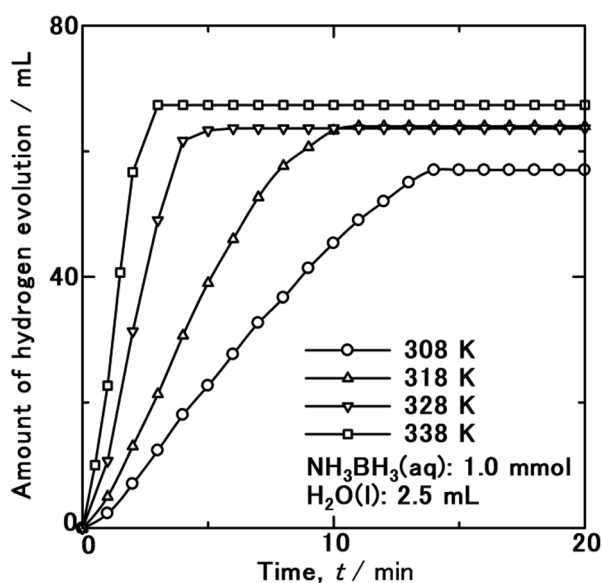


Fig. 3 HEV from hydrolysis of NH_3BH_3 aqueous solution (1 mmol, 2.5 ml) using WC- $\text{Co}_{\text{carbide}}$ as a catalyst at various temperatures.

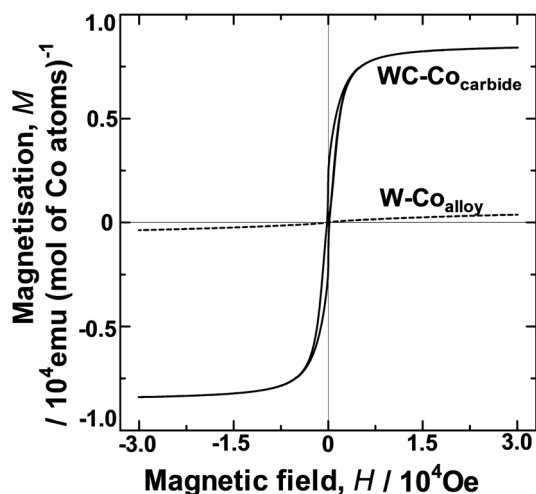


Fig. 5 Magnetisation as a function of magnetic field for WC-Co_{carbide} and W-Co_{alloy} at 4 K.

value of M for WC-Co_{carbide} resulted from the Co crystals, the maximum saturated value, M_s , was 8.41×10^3 emu (mol Co atoms)⁻¹. However, the value of M for W-Co_{alloy} was close to zero, indicating that it had a spin glass state.³⁰ Similarly, the M (H) data were measured at 308 K, yielding similar results to those at 4 K for both samples.

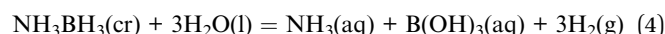
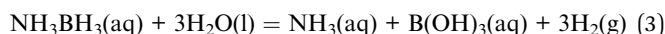
The magnetic moment β of a Co atom in the Co nanocrystals in WC-Co_{carbide} at 4 K and 308 K were 1.50 and 1.42 μ_B per Co atom, respectively, where μ_B is the unit of the Bohr magneton. In the elemental Co crystal, the value of β for a Co atom is 1.7 μ_B .²⁷ Hence, the Co nanocrystals had a β value close to that of elemental Co. In contrast, the β value of a Co atom substituted into the bcc W lattice in the W-Co_{alloy} measured at 4 K and 308 K was 6.79×10^{-2} and 7.48×10^{-3} μ_B , respectively. As described in our previous study,²⁴ the XRD pattern of W-Co_{alloy} showed only the peaks of W.²⁴ The composition of W-Co_{alloy} corresponds to the two phase equilibria region of the W phase and the intermediate phase of Co₇W₆ in the phase diagram of the W-Co binary system.³¹ Co₇W₆ has complicated long-range periodic structure.³² Formation of the complicated long-range periodic structures such as the sigma phase in the heat-resistant alloys takes a lot of time to accomplish their atomic configuration.³³ Hence, the equilibrium phase of Co₇W₆ was not formed during reduction with H₂ gas. The Co atoms are concluded to be supersaturated in the W lattice as a nonequilibrium state. Consequently, the spin configuration among isolated Co atoms was random (*i.e.*, a spin glass state³⁰). Hence, the spin-ordered state of the Co in WC-Co_{carbide} appears to be one of the factors determining the singular catalytic activity.

Cerri *et al.*¹⁷ found that hydrogen chemisorption induced a disordering of electronic spin polarisation on the surface of ferromagnetic Gd, resulting in a change in M , and the Curie temperature, T_C . It is likely that hydrogen chemisorption relaxes the spin configuration in the system. In aqueous solutions, the spin configuration of the protons is relevant. Here, the protons are likely to be absorbed on the ferromagnetic Co nanocrystals,

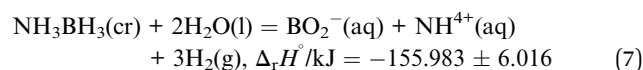
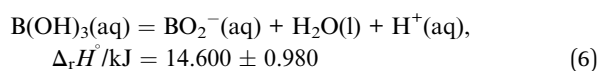
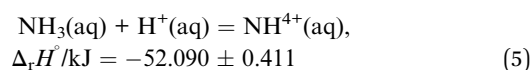
such that their nucleus spin configurations are aligned to be antiparallel to relax the spin polarisation of the surface.

Thermodynamic cycle

The following equations show the thermodynamic equilibria among the species associated with the HER from hydrolysis of NH₃BH₃(cr), where the standard enthalpies of formation, $\Delta_f H_m^\circ$, at 298.15 K of the standard substances of NH₃BH₃(cr),³⁴ H₂O(l),³⁵ ammonia (NH₃(aq)),³⁵ orthoboric acid (B(OH)₃(aq)),³⁵ ammonium (NH₄⁺(aq)),³⁵ metaboric acid (BO₂⁻(aq)),³⁶ and H₂(g)³⁵ are summarised in Table S1.† Eqn (2) shows the hydration reaction of NH₃BH₃(cr), where the thermodynamic value is unknown. Eqn (3) shows the HER of the hydrolysis of NH₃-BH₃(aq). Eqn (4), rewritten as the sum of eqn (2) and (3), indicates the HER from the initial substance of NH₃BH₃(cr). Eqn (5) shows the formation of NH₄⁺(aq). Eqn (6) shows the formation of BO₂⁻(aq). Finally, eqn (7), rewritten as the sum of eqn (4)–(6), shows the final state of the hydrolysis.



$$\begin{aligned} \Delta_r H^\circ / \text{kJ} = & 3 \times \Delta_f H_m^\circ(\text{H}_2(\text{g})) + \Delta_f H_m^\circ(\text{NH}_3(\text{aq})) \\ & + \Delta_f H_m^\circ(\text{B}(\text{OH})_3(\text{aq})) - \Delta_f H_m^\circ(\text{NH}_3\text{BH}_3(\text{cr})) \\ & - 3 \times \Delta_f H_m^\circ(\text{H}_2(\text{g})) = -118.494 \pm 5.921 \end{aligned}$$



Since eqn (5)–(7) are spontaneous reactions, the HER is given by eqn (4). As the standard entropy, S_m° , of NH₃BH₃(cr) has not yet been measured, the standard entropy of reaction, $\Delta_r S^\circ$, and the standard Gibbs energy of reaction, $\Delta_r G^\circ$, are unknown. However, $\Delta_r G^\circ$ is more negative than $\Delta_r H^\circ$ as the HER increases $\Delta_r S^\circ$. Therefore, when a driving energy is applied corresponding to the hydrogen overpotential of Co, the HER reaches equilibrium, as defined by eqn (7) *via* eqn (4).

In previous studies, Co NPs (10 wt% Co on an Al₂O₃ support)²¹ had an HER rate 10 times lower than that of Pt NPs (2 wt% Pt on an Al₂O₃ support).²² However, in the present study, the HER rate of WC-Co_{carbide} was 30% higher than that of Pt NPs, even if E_a corresponds to the η_{H_2} of Co. The excellent catalytic activity was a result of the high A value, as determined from the Arrhenius plot (Fig. 4). The WC matrix seems to facilitate the release of H⁺(aq) from NH₃BH₃(aq) and contribute to increasing A (*i.e.*, the number of H⁺(aq) collisions). In accordance with the electron theory, the Pt-like high DOS near

the E_F of the surface of WC⁷ can induce adsorption of NH_3BH_3 molecules, which can promote the decomposition of B–N bonds to form $\text{NH}_3(\text{aq})$, $\text{B}(\text{OH})_3(\text{aq})$, and $\text{H}_2(\text{g})$ via highly unstable $\text{BH}_3(\text{aq})$ close to the equilibrium states (see eqn (3)). Furthermore, considering the thermodynamic hierarchy, as shown in Table S1,† the $\Delta_f G_m^\circ$ of WB³⁷ is smaller than that of WC,³⁸ indicating that the chemical bonding between the W and B atoms is more stable than that between the W and C atoms. In addition, it is well known that natural tungsten ore consists of ammonium tungstate,³⁹ which is the starting material used in the present study, indicating the high affinity between tungsten and ammonia. Therefore, W in the WC matrix has a driving force for attracting $\text{NH}_3\text{BH}_3(\text{aq})$, resulting in preferential adsorption. Although a previous study²⁰ investigated the hydrolysis kinetics of NH_3BH_3 using first principles calculations based on the transition state theory and estimated the atomic distance of a B–N bond, the interaction between the WC matrix and $\text{NH}_3\text{BH}_3(\text{aq})$ should be further investigated by first principles and molecular dynamics calculations. The $\text{BH}_3(\text{aq})$ molecule can release three protons while bonding with three OH^- ions to form $\text{B}(\text{OH})_3$. As shown in Table 1, when the amount of $\text{NH}_3\text{BH}_3(\text{aq})$ was doubled, N_{HER} also doubled, indicating that there were sufficient HER reaction sites. Hence, the WC matrix played a crucial role in adsorbing NH_3BH_3 molecules and decomposing B–N bonds, followed by supplying protons to the Co crystals.

Specific enthalpy of hydrogen adsorption

The $\Delta_{\text{ad}} H_m^\circ$ values of WC–Co_{carbide}, and commercial WC and Pt powders at 423 K were -22.42 ± 0.90 , -21.87 ± 0.87 , and $-109.72 \pm 4.39 \text{ J m}^{-2}$, respectively. We found that the $\Delta_{\text{ad}} H_m^\circ$ of WC was less exothermic than that of Pt, implying that the desorption of $\text{H}_2(\text{g})$ from the surface of WC was not difficult. Previous studies have investigated the catalytic activity of WC¹³ and related carbide composites^{11,12} under an external voltage. When an external voltage is applied, the $\text{H}^+(\text{aq})$ ions align on the surface of WC, followed by hydrogen evolution. However, when no external voltage is applied, such an alignment of $\text{H}^+(\text{aq})$ is unlikely to be significant. In contrast, on the surface of WC–Co_{carbide}, the $\text{H}^+(\text{aq})$ ions seem to align as a result of the magnetic field of the Co nanocrystals, followed by hydrogen evolution. The $\Delta_{\text{ad}} H_m^\circ$ of Pt is extremely high, indicating that hydrogen was stored in its lattice. Pt simultaneously achieves the adsorption, storage, and desorption of hydrogen species, although the specific mechanism is still unknown.

Mechanism of a singular catalytic behaviour of WC–Co_{carbide}

Fig. 6 shows a schematic of the catalytic HER over WC–Co_{carbide}, which involves the following steps. (1) The NH_3BH_3 molecules are absorbed on the WC matrix by the attractive interaction of W–B bonds; (2) the B–N bonds in $\text{NH}_3\text{BH}_3(\text{aq})$ are broken to form stable $\text{NH}_3(\text{aq})$ and unstable $\text{BH}_3(\text{aq})$; (3) the B atoms in $\text{BH}_3(\text{aq})$ with a significantly short lifetime release three quasi-stable protons and coordinate three $\text{OH}^-(\text{aq})$ ions from the surrounding H_2O molecules to form stable $\text{B}(\text{OH})_3$ molecules, wherein the H_2O releases three protons; (4) sufficient amounts

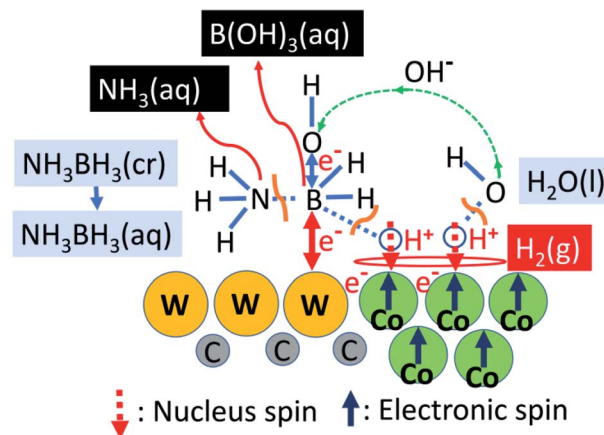


Fig. 6 Schematic of hydrogen evolution from the hydrolysis of $\text{NH}_3\text{BH}_3(\text{aq})$ over the WC–Co_{carbide} catalyst.

of protons from $\text{BH}_3(\text{aq})$ and H_2O are adsorbed on the ferromagnetic Co nanocrystals in the WC matrix, resulting in anti-parallel alignment of the nucleus spin configurations of the two $\text{H}^+(\text{aq})$ to relax the spin polarisation of the system; (5) hydrogen molecules are evolved as the protons accept electrons, e^- , such that E_a is consistent with η_{H_2} of Co, wherein Co is protected by galvanic protection;⁴⁰ (6) one e^- is released during the decomposition of BH_3 , along with a proton, while the other e^- is released from $\text{OH}^-(\text{aq})$ during coordination to form $\text{B}(\text{OH})_3(\text{aq})$, i.e. the origin of charge transfer is the B atom adsorbed on the W atom, wherein the charge is transferred to Co via the WC matrix; (7) the WC matrix decomposing $\text{NH}_3\text{BH}_3(\text{aq})$ and generating $\text{H}^+(\text{aq})$ applies a driving energy corresponding to the necessary hydrogen overpotential, wherein the WC matrix supplies excess protons to the Co sites.

The relative volume ratio of the WC matrix vs. Co nanocrystals is estimated to be 100 : 13 from the densities of pristine WC and Co. When the relative SSA ratio of the WC matrix vs. Co nanocrystals is hypothetically equal to the relative volume ratio, the SSA of the Co nanocrystals is $0.15 \text{ m}^2 \text{ g}_{\text{cat}}^{-1}$. Assuming the Co nanocrystals only contribute HER, N_{HER} is evaluated as $33.84 (\text{H}_2 \text{ mol min}^{-1} \text{ m}^{-2})$ which is twelve times faster than that of Pt NPs ($=2.90 (\text{H}_2 \text{ mol min}^{-1} \text{ m}^{-2})$). Such an extremely fast N_{HER} is never caused by the single Co nanocrystals. Hence, the Pt-like catalytic activity is attributed to the synergistic effect of the WC matrix and ferromagnetic Co nanocrystals.

Previous studies have determined the E_a values for the HER in $\text{NH}_3\text{BH}_3(\text{aq})$ of 21 kJ mol^{-1} over Pt NPs²² and 62 kJ mol^{-1} over Co NPs,²¹ and suggested that the rate determining step (RDS) was the cleavage of the B–N bonds, as unstable BH_3 reacts with H_2O to form H_2 . In addition, Wang *et al.*²³ determined an E_a value of 42.7 kJ mol^{-1} for the HER in $\text{NH}_3\text{BH}_3(\text{aq})$ catalysed by Ni NPs, and suggested that the RDS was the cleavage of the O–H bond in H_2O ($E_a = 493 \text{ kJ mol}^{-1}$), as the equilibrium bonding energy was more endothermic than that of the B–N ($E_a = 117 \text{ kJ mol}^{-1}$) and B–H (430 kJ mol^{-1}) bonds. The ΔE_a between Co NPs²¹ and Pt NPs²² and that between Ni NPs²³ and Pt NPs²² were 41 and 21.7 kJ mol^{-1} , corresponding to -0.43 and -0.22 V , respectively, thereby demonstrating consistency with the η_{H_2} of

Co and Ni.²⁹ These E_a results support the mechanism depicted in Fig. 6.

Another definition of η_{H_2} in electrochemistry is the HER current density, i_0 , determined by extrapolating the cathodic Tafel line to the reference hydrogen electrode.⁴¹ Fig. S4† shows the correlation between magnetic susceptibility, X ,⁴² and $\log i_0$ ⁴¹ of the transition metals of the fourth, fifth, and sixth periods (rows) of the periodic table of the elements. In the sixth period, the $\log i_0$ value of Pt was large, consistent with the high HER catalytic activity, while X was large, corresponding to a high $\log i_0$. Similarly, in the fifth period, the $\log i_0$ values and X values of Rh and Pd were large. In the fourth period, the $\log i_0$ values of Fe, Co, and Ni were high. As these metals are ferromagnetic with spontaneous magnetisation, their X values are not relevant. However, their spin-ordered ferromagnetic state likely increased their $\log i_0$ values. In contrast, the diamagnetic transition metals of Au and Hg in the sixth period, Ag and Cd in the fifth period, and Cu and Zn in the fourth period with negative X values showed small $\log i_0$ values consistent with their lower HER catalytic activities. Therefore, we conclude that the magnetic properties of the transition metals are related to the HER catalytic activity. Materials with high X values are likely to interact readily with the nucleus spin of $H^+(aq)$. Ferromagnetism seems to induce alignment of the nucleus spin of $H^+(aq)$ on the ferromagnetic metal surface, as discussed with regard to the contribution of the Co nanocrystals in this study. After formation of two radical H atoms by donation of electrons, the spin conversion of the two H atoms should form the $1\sigma_g$ molecular orbital of a H_2 molecule. The magnetic properties of metals are likely to contribute to the spin conversion. After H_2 molecules are formed, they are released from the surface of the metals with a high positive X value or ferromagnetism because H_2 is diamagnetic with a negative X value ($-19.7 \times 10^{-9} \text{ m}^3 \text{ kg}^{-1}$).

Recently, it was shown that an external magnetic field enhanced the oxygen evolution reaction (OER) during electrolysis of a KOH aqueous solution.¹⁹ A Pt plate was used as the cathode and ferrite $NiZnFe_4O_x$ (deposited on Ni foam by direct magnetic interaction between the materials) was used as the anode. An external magnetic field of 0.45 T was applied to the anode using a rare-earth permanent magnet, which resulted in the OER current almost doubling (from 24 to 40 mA cm^{-2} at $\geq 1.65 \text{ V}$).¹⁹ They hypothesised that the external voltage accelerated the spin conversion to form the paramagnetic triplet of $O_2(g)$. Li *et al.*^{43,44} prepared a series of metal–organic frameworks (MOFs) composed of the Fe–Ni binary⁴³ and W–Co–Fe ternary system⁴⁴ as electrocatalysts. Their superior catalytic properties appear to result from the magnetic elements of Fe,^{43,44} Ni⁴³, and Co.^{43,44} These findings are highly relevant to the present study.

There have been many investigations of the electronic states of HER.²⁹ However, the interaction between the nucleus spin of $H^+(aq)$ and electronic spin of metals has not yet been investigated. Thomas *et al.*^{45,46} calculated the energy eigenvalues of the electrons and nuclei of hydrogen atoms (protons) by considering the wave functions of both protons and electrons, and describing the zero-point motion of the protons by a Slater-type

wave function. This approach could be applied to the HER catalytic reaction.

As described above with respect to the thermodynamics, the WC matrix appears to facilitate the release of $H^+(aq)$ from $NH_3BH_3(aq)$ and contribute to increasing A (*i.e.*, the number of collisions between $H^+(aq)$ ions). However, $H_2(g)$ appears to be barely formed by such collisions. The alignment of $H^+(aq)$ on the catalyst surface and the donation of an electron to the ion are necessary. When an external electric voltage is applied to WC, these requirements are likely to be satisfied, resulting in the catalytic activity observed in previous studies.^{11–16} The internal magnetic field is considered sufficient to induce catalytic activity without requiring an external voltage.

Pt seems to easily adsorb $NH_3BH_3(aq)$ with the highest DOS at the E_F of all elements, and easily evolves $H_2(g)$ from its surface with the smallest η_{H_2} of any known element. A study of the interaction between Pt with a high magnetic susceptibility and protons with a nucleus spin appears to be a necessary future work to consider the effect of the magnetic field from the earth on the spin alignment of the protons.

Pt catalysts have low reusability in HER from HCOOH due to CO poisoning. WC–Co_{carbide} is advantageous for avoiding CO poisoning due to that it is prepared by carburizing with CO *via* CO₂. Reusability of WC–Co_{carbide} should be further investigated.

Conclusion

The catalytic activities of WC and its composites had previously only been observed under the conditions of an applied external voltage, and the predicted intrinsic catalytic activity of mono WC without an external voltage had not been verified. We hypothesised that the introduction of an internal magnetic field would provide conditions similar to those provided by the external voltage to enhance the catalytic performance of WC. To this end, the WC lattice was doped with ferromagnetic Co nanocrystals to introduce an ordered-spin configuration as an internal magnetic field. The internal magnetic field successfully increased the HER rate during hydrolysis of NH_3BH_3 to a value even higher than that of the Pt nanoparticles. The enhanced catalytic activity was attributed to the synergistic effect of the WC matrix promoting hydrolytic cleavage of NH_3BH_3 and the ferromagnetic Co crystals interacting with the nucleus spin of the protons. Our new strategy for enhancing the catalytic performance by introducing an internal magnetic field is expected to provide new opportunities for developing novel catalysts, such as those for hydrolysis of NH_3BH_3 and the electrodes of fuel cells.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was partially supported by an academic research grant project provided by the Hyogo Science and Technology Association.



Notes and references

- 1 R. B. Levy and M. Boudart, *Science*, 1973, **181**, 547.
- 2 L. H. Bennett, J. R. Cuthill, A. J. McAlister and N. E. Erickson, *Science*, 1974, **184**, 563.
- 3 J. B. Ketterson, F. M. Mueller and L. R. Windmiller, *Phys. Rev.*, 1969, **186**, 656.
- 4 L. F. Mattheiss, *Phys. Rev.*, 1965, **139**, 1893.
- 5 J. E. Houston, G. E. Laramore and R. L. Park, *Science*, 1974, **185**, 258.
- 6 L. H. Bennett, J. R. Cuthill, A. J. McAlister and N. E. Erickson, *Science*, 1975, **187**, 858.
- 7 L. F. Mattheiss and D. R. Hamann, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1984, **15**, 1731.
- 8 P. Blaha, K. Schwarz, F. Tran, R. Laskowski, G. K. H. Madsen and L. D. Marks, *J. Chem. Phys.*, 2020, **152**, 074101.
- 9 X. Zhang, B. Grabowski, C. Freysoldt, F. Körmann and J. Neugebauer, *Phys. Rev. B*, 2017, **95**, 165126.
- 10 P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 11 D. V. Esposito, S. T. Hunt, A. L. Stottlmyer, K. D. Dobson, B. E. McCandless, R. W. Birkmire and J. G. Chen, *Angew. Chem., Int. Ed.*, 2010, **49**, 9859.
- 12 J. Zhang, J. Chen, Y. Jiang, F. Zhou, G. Wang, R. Wang and R. Wang, *Appl. Surf. Sci.*, 2016, **389**, 157.
- 13 D. J. Ham, R. Ganesan and J. S. Lee, *Int. J. Hydrogen Energy*, 2008, **33**, 6865.
- 14 C. Ma, J. Sheng, N. Brandon, C. Zhang and G. Li, *Int. J. Hydrogen Energy*, 2007, **32**, 2824.
- 15 P. Xiao, X. Ge, H. Wang, A. Liu, A. Fisher and X. Wang, *Adv. Funct. Mater.*, 2015, **25**, 1520.
- 16 W. Zheng, L. Wang, F. Deng, S. A. Stephen, A. Gilles, A. K. Prasad, S. G. Advani, Y. Yan and D. G. Vlachos, *Nat. Commun.*, 2017, **8**, 418.
- 17 A. Cerri, D. Mauri and M. Landolt, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1983, **27**, 6526.
- 18 T. Das, S. Kweon, J. Choi, S. Y. Kim and I. Oh, *Int. J. Hydrogen Energy*, 2015, **40**, 383.
- 19 F. A. Garcés-Pineda, M. Blasco-Ahicart, D. Nieto-Castro, N. Lopez and J. R. Galan-Mascalos, *Nat. Energy*, 2019, **4**, 519.
- 20 T. Banu, T. Debnath, T. Ash and A. K. Das, *J. Chem. Phys.*, 2015, **143**, 194305.
- 21 Q. Xu and M. Chandra, *J. Power Sources*, 2006, **163**, 364.
- 22 M. Chandra and Q. Xu, *J. Power Sources*, 2007, **168**, 135.
- 23 C. Wang, J. Tuninetti, Z. Wang, C. Zheng, R. Ciganda, L. Salmon, S. Moya, J. Ruiz and D. Astruc, *J. Am. Chem. Soc.*, 2017, **139**, 11610.
- 24 M. Morishita, H. Yamamoto, T. Okahira, M. Yoshioka, N. Fukumuro, H. Matsuda, M. Ikebe, M. Iwasaki, H. Yanagida and H. Nishimaki, *J. Am. Ceram. Soc.*, 2012, **95**, 3797.
- 25 K. Nakai and K. Nakamura, *Proceedings of the 22nd reference catalyst forum*, 2001, p. 31.
- 26 J. R. Taylor, *An Introduction to Error Analysis: The Study of Uncertainty in Physical Measurement*, Oxford University Press, Oxford, 1982, p. 57.
- 27 S. Chikazumi, *Physics of Ferromagnetism Vol. 1: Magnetic Properties of Matter*, Shokabo, Tokyo, 22nd edn, 2007, pp. 7, 11, 47, 184.
- 28 C. Kittel, *Introduction to Solid State Physics*, John Wiley & Sons, Inc., New York, 6th edn, 1986, p. 426.
- 29 H. Ezaki, M. Morinaga and S. Watanabe, *Electrochim. Acta*, 1993, **38**, 557.
- 30 V. Cannella and J. A. Mydosh, *Phys. Rev. B: Solid State*, 1972, **6**, 4220.
- 31 S. V. Nagender Naidu, A. M. Sriramamurthy and P. Rama Rao, Co-W, in *Binary Alloy Phase Diagrams*, ed. T. B. Massalski, ASM International: The Materials Information Society, Materials Park, OH, 2nd edn, 1990, pp. 1257–1259.
- 32 P. Villars and L. D. Calvert, Co₇W₆, in *Peason's Handbook of Crystallographic Data for Intermetallic Phases*, ed. P. Villars and L. D. Calvert, ASM, Materials Park, OH, 1985, vol. 2, p. 1845.
- 33 M. Morinaga, N. Yukawa and H. Ezaki, Solid Solubilities in Transition Metal-Based f.c.c. Alloys, *Philos. Mag. A*, 1985, **51**, 223.
- 34 Y. K. Shaulov, G. O. Shmyreva and S. Tubyanskaya, *Zh. Fiz. Khim.*, 1966, **40**, 122.
- 35 H. Gamsjäger, J. Bugajski, T. Gajda, R. J. Lemire and W. Preis, Chemical Thermodynamics of Nickel, in *Chemical Thermodynamics*, Nuclear Energy Agency, Organization for Economic Co-operation, Development, Elsevier Science Publisher B.V., Amsterdam, 2005, vol. 6, p. 43.
- 36 D. D. Wagman, W. H. Evans, V. B. Parker, R. H. Schumm, I. Halow, S. M. Bailey, K. L. Churney and R. L. Nuttall, *J. Phys. Chem.*, 1982, **11**, 123.
- 37 H. Yamamoto, M. Morishita and S. Hiramatsu, *Metall. Mater. Trans. B*, 2017, **48**, 1703.
- 38 I. Barlin, *Thermochemical Data of Pure Substance*, VCH Verlag, part 1, 1989, p. 1642.
- 39 T. Azakami, Y. Awakura and M. Itagaki, *Extractive Metallurgy*, Japan Inst. Met., Maruzen, Tokyo, 1999, p. 167.
- 40 M. Morishita, K. Koyama, M. Murase and Y. Mori, *ISIJ Int.*, 1996, **36**, 714.
- 41 H. Kita and K. Uozaki, *Introduction to Electrochemistry*, Gihoudo, Tokyo, 1987, p. 239.
- 42 *Metals Data Book*, ed. S. Nagasaki, Japan Inst. Met., Maruzen, Tokyo, 1993, pp. 18–19.
- 43 Y. Li, M. Lu, Y. Wu, Q. Ji, H. Xu, J. Gao, G. Qian and Q. Zhang, *J. Mater. Chem. A*, 2020, **8**, 18215.
- 44 Y. Li, T. Zhao, M. Lu, Y. Wu, Y. Xie, H. Xu, J. Gao, J. Yao, G. Qian and Q. Zhang, *Small*, 2019, **15**, 1901940.
- 45 I. L. Thomas, *Phys. Rev.*, 1969, **185**, 90.
- 46 I. L. Thomas and H. W. Joy, *Phys. Rev. A*, 1970, **2**, 1200.

