

Cite this: *J. Mater. Chem. C*, 2015,
3, 5940

The effect of regioisomerism on the solid-state fluorescence of bis(piperidyl)anthracenes: structurally simple but bright AIE luminogens†

Shunsuke Sasaki,^a Kazunobu Igawa^b and Gen-ichi Konishi*^{ac}

A series of regioisomers of piperidylanthracenes (PA) and bis(piperidyl)anthracenes (BPA) were synthesized and their photophysical properties examined in solution, suspension, and the solid state. Aggregation-induced emission (AIE) was observed only for **1,4-BPA** and **9,10-BPA**, in which two piperidyl groups are substituted at the anthracene moiety in the *para*-position with respect to each other. Compared to previously reported AIE luminogens, these easily obtainable *para*-substituted BPAs, exhibited several unique and beneficial features, such as simple structures, bright solid-state fluorescence ($\Phi_{fl} = 0.49$ and 0.86 for **1,4-BPA** and **9,10-BPA**, respectively), tunable fluorescence emission, and large Stokes shifts. Results from diffuse-reflectance and fluorescence lifetime measurements demonstrated that PAs and BPAs intrinsically possess an efficient non-radiative transition pathway in the solid state, and that **1,4-BPA** and **9,10-BPA** may overcome this pathway. The X-ray crystallographic analysis of **1,4-BPA** revealed that undesirable interchromophoric interactions can be minimised, while TD-DFT calculations suggested that the enhanced Stokes shift of **1,4-BPA** arises from severe electronic repulsion between neighbouring piperidyl moieties, which presumably results in the absence of self-absorption.

Received 4th April 2015,
Accepted 11th May 2015

DOI: 10.1039/c5tc00946d

www.rsc.org/MaterialsC

Introduction

Most organic fluorophores suffer from substantial fluorescence quenching in the solid state, an effect usually referred to as aggregation-caused quenching (ACQ). For decades, this characteristic feature has often been considered a bottleneck for practical applications of such organic compounds in *e.g.* electronics and bioimaging devices.^{1,2} However, traditional concepts for organic fluorophores have recently changed dramatically, as various studies have described simple and yet accurate systems for the measurement of absolute fluorescence quantum yields (Φ_{fl}).^{3,4} The fluorescence quantum yield is generally considered a quantitative index for solid-state fluorescence,⁵ and it provides deeper insight into possible new design strategies for efficient solid-state fluorophores, thereby promoting the development of fluorophores that exhibit aggregation-induced

emission (AIE)^{1,2} and efficient solid-state fluorescence.^{6,7} Especially AIE luminogens have found many applications, not only in organic light-emitting diodes (OLEDs),^{8,9} but also in environment-responsive fluorescent probes,^{1,10} which are sensitive to aggregation and self-assembly.

Despite the development of a wide variety of AIE luminogens^{1,2} and efficient solid-state fluorophores,^{6,7} the majority of functional molecules used in OLEDs^{8,9} and fluorescence imaging¹¹ are based on tetraphenylethenes (TPEs),¹² stilbenes,^{13,14} and siloles.¹⁵ These chromophores are preferentially used, because they exhibit simple structures and are susceptible to further modifications, which may be necessary for certain applications.¹⁶ Conversely, most other fluorophores exhibit more complex and complicated structures. Most recently, Tang *et al.* have addressed this issue and developed simple, small, and easily accessible AIE luminogens based on tetraphenylpyrazines (TPPs).¹⁷ Although both TPE and TPP are more readily available than other AIE luminogens, they exhibit Φ_{fl} values that remain unsatisfactory for most applications^{12,17} (Chart 1), and further modifications are required in order to attain efficient solid-state fluorescence.¹⁸

In order to develop simple, easily accessible, and versatile AIE luminogens, we focused on the solid-state fluorescence behaviour of basic dyes.¹⁹ Bis(dialkylamino)anthracene is a promising candidate in this context, as it does not exhibit

^a Department of Organic and Polymeric Materials, Tokyo Institute of Technology, Tokyo 152-8552, Japan. E-mail: konishi.g.aa@m.titech.ac.jp; Fax: +81-3-5734-2321

^b Institute for Materials Chemistry and Engineering and Department of Molecular and Material Sciences, Kyushu University, Fukuoka 816-8580, Japan

^c PRESTO, Japan Science and Technology Agency (JST), Japan

† Electronic supplementary information (ESI) available: Spectral data, quantum chemical calculations, and X-ray crystal structure analyses. CCDC 1055226–1055229. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5tc00946d



| | TPE | 9,10-BPA | 1,4-BPA |
|-------------------------------|---|--|---------------|
| Φ_{fl} (solution) | 0.002 (Acetonitrile) | 0.02 (DMF) | 0.08 (DMF) |
| Φ_{fl} (solid) | 0.15 (Aggregate) | 0.86 (Powder) | 0.49 (Powder) |
| Colour tuning | Extension of the conjugation / Introduction functional groups | Regioisomerism of the piperidyl groups | |
| Preparation | One step from benzophenone | One step from dibromoanthracene | |

Chart 1 Comparison between TPE¹² and AIE-active BPAs.

Chart 2 Nomenclature of piperidyl-substituted PA and BPA derivatives used in this study.

fluorescence in solution.²⁰ This is probably due to fast internal conversion²¹ and can be suppressed by aggregation. However, it is exactly this fluorescence property that sets bis(dialkylamino)anthracene apart from other fluorescent analogues such as *e.g.* diphenylanthracene.³

In this study, we report the synthesis of various regioisomers of piperidylanthracenes (PAs) and bis(piperidyl)anthracenes (BPAs), and we discuss the effect of regioisomerism on their solid-state fluorescence properties (Chart 2). Our results revealed that only **1,4-BPA** and **9,10-BPA** exhibit AIE behaviour, and that their fluorescence colours are tuneable without the introduction of additional substituents (Chart 1). Moreover, **9,10-BPA**, which can be easily obtained *via* a one-step reaction from inexpensive 9,10-dibromoanthracene, also exhibits highly efficient solid-state fluorescence ($\Phi_{\text{fl}} = 0.86$; neat powder). As BPAs are prepared with relative ease, and as they exhibit considerable levels of AIE, efficient solid-state fluorescence, characteristic tunability of the fluorescence colour, as well as simple and small structures, they are ideally suited to serve as core fluorophores for various functional molecules.

However, BPAs are not only attractive with respect to the aforementioned practical aspects, but they are also interesting from a fundamental perspective, as they display efficient solid-state fluorescence as well as a large fluctuation in their Stokes shifts. This is particularly remarkable, as the molecular mass of BPAs is relatively low and they do not contain aryl–aryl bonds, alkene moieties, or bulky substituents.^{1,2,6} Therefore, the fluorescence behaviour of BPAs will also be discussed from a mechanistic viewpoint.

Results and discussion

Synthesis

A series of PA and BPA derivatives was readily obtained from the Pd-catalysed C–N coupling of mono- and dibromoanthracene,

respectively.^{22,23} This route was given preference over a previously reported approach starting from 9,10-bis(dimethylamino)anthracene,^{20,24} as the synthetic intermediate 9,10-diaminoanthracene generated therein is unstable and susceptible to rapid oxidation.²⁵ Most likely, this instability is responsible for the shortage of reports on the emission properties of bis(dialkylamino)anthracenes.^{26–30} Fortunately, recent advances in Pd-catalysed C–N coupling reactions provide synthetic routes to PAs and BPAs that bypass the formation of this unstable intermediate. Furthermore, **1,4-BPA** and **9,10-BPA** thus prepared do not exhibit any signs of deterioration upon exposure to air, which was attributed to the highly twisted aryl–nitrogen bonds, which suppresses oxidation *via* the formation of a quinoidal intermediate.²⁵ Thermogravimetric analysis (TGA) measurements revealed decomposition temperatures (T_{d} , corresponding to 5% weight loss under a nitrogen atmosphere) of 245 and 253 °C for **1,4-BPA** and **9,10-BPA**, respectively (Fig. S17, ESI†). These T_{d} values are higher than that of **TPE** ($T_{\text{d}} = 213$ °C) and comparable to that of **TPP** ($T_{\text{d}} = 275$ °C),¹⁷ suggesting that **1,4-BPA** and **9,10-BPA** should possess a sufficiently high thermal stability in order to serve as AIE luminogens. Detailed synthetic procedures and characterization data are summarized in the experimental section.

Aggregation-induced emission (AIE)

We measured the fluorescence properties of PAs and BPAs in solution (DMF), in colloidal suspension (THF : H₂O = 1 : 9; v/v), and in the solid state (neat powder, Fig. 1). The formation of aggregates was supported by UV-vis spectroscopy^{19,31} and confirmed by dynamic light scattering (DLS) measurements. For PA and BPA, the DLS measurements revealed typical hydrodynamic radii (R_{h}) of *ca.* 50–100 nm (Fig. S19–S26, ESI†). More detailed information about these DLS measurements is contained in the experimental section. Table 1 summarises the fluorescence maxima (λ_{fl}) and fluorescence quantum yields (Φ_{fl}) for each state. Most PAs and BPAs suffer severely from ACQ, and only **1,4-BPA** and **9,10-BPA**, which contain piperidyl groups at the *para*-position of the anthracene moiety with respect to each other, exhibit AIE behaviour. Especially **9,10-BPA** demonstrates noteworthy levels of AIE, as it lacks fluorescence in solution ($\Phi_{\text{fl}} \approx 0.02$), but exhibits efficient solid-state fluorescence ($\Phi_{\text{fl}} \approx 0.86$), which is comparable to previously reported efficient fluorophores.^{6,7,18,32,33} Interestingly, **1,4-BPA** reduces its fluorescence quantum yield as a function of increasing solvent polarity, and significantly enhanced fluorescence is

Fig. 1 Photographic images of BPAs exposed to (a) ambient light and (b) UV light ($\lambda = 365$ nm).

Table 1 Fluorescence maxima (λ_{fl}) and fluorescence quantum yield (Φ_{fl}) for **PAs** and **BPA**s in solution, suspension, and the solid state

| Entry | λ_{fl}^a [nm] | | | | Φ_{fl}^a | | | |
|-----------------|-----------------------|-----|-------------------------|-----|---------------|------|--------------------|------|
| | Solution | | Suspension ^b | | Solution | | Solid ^c | |
| 1-PA | 521 | 550 | 519 | 500 | 0.91 | 0.85 | 0.49 | 0.52 |
| 2-PA | 504 | 525 | 488 | 523 | 0.91 | 0.92 | 0.34 | 0.27 |
| 9-PA | 478 | 495 | 490 | 519 | 0.01 | 0.01 | 0.05 | 0.71 |
| 1,4-BPA | 598 | 631 | 596 | 591 | 0.25 | 0.08 | 0.45 | 0.49 |
| 1,5-BPA | 504 | 533 | 504 | 518 | 0.77 | 0.87 | 0.12 | 0.34 |
| 1,8-BPA | 514 | 540 | 514 | 531 | 0.71 | 0.80 | 0.15 | 0.12 |
| 2,6-BPA | 515 | 528 | 504 | 532 | 0.88 | 0.92 | 0.15 | 0.17 |
| 9,10-BPA | 528 | 537 | 514 | 519 | 0.02 | 0.02 | 0.79 | 0.86 |

^a The excitation wavelength corresponds to either the maximum wavelength of the absorption spectrum, or the diffused-reflectance spectrum. ^b THF:H₂O = 1:9 (v/v); H₂O was added in order to obtain colloidal suspensions, and the formation of aggregates was confirmed by UV-vis spectroscopy^{19,31} and DLS measurements. ^c Data obtained from neat powder.

observed in colloidal suspension (Table 1). This polarity-reduced fluorescence and the AIE characteristics are clearly reflected in the fluorescence behaviour when water is added to a THF solution of **1,4-BPA** (Fig. 2a and c). Then, the fluorescence intensity plot is quite similar to AIE luminogens, whose fluorescence is quenched substantially by the addition of polar solvents.³⁴ Moreover, the fluorescence colour of **PAs** and **BPA**s varies, depending on the position of the substituents. For example, green fluorescence (λ_{fl} = 514–519 nm) was observed for **9,10-BPA**, while orange fluorescence (λ_{fl} = 591–596 nm) was observed for **1,4-BPA**. Such a drastic colour change induced by regioisomerism is quite unique compared to conventional

design paradigms, where the fluorescence colour is modulated *via* the introduction of donor or acceptor groups.³³ This intriguing feature may be applied to the multicolour imaging of living cells, which imposes strict uniformity requirements on the chemical structures of the fluorophores.^{11,35} In addition, large Stokes shifts were observed for both **1,4-BPA** (7400 cm⁻¹) and **9,10-BPA** (5500 cm⁻¹) in the solid state, which renders these compounds attractive for applications in far-field fluorescence microscopy.³⁶ In fact, these Stokes shifts are at least comparable, if not larger than those reported for recently developed fluorescent probes with large Stokes shifts.³⁷

Effects of aggregation on the radiative and non-radiative transition pathways of **PAs** and **BPA**s

In contrast to the other regioisomers, only *para*-substituted **BPA**s such as **1,4-BPA** and **9,10-BPA** exhibit AIE. In order to understand this unusual phenomenon, the correlation between aggregation (or crystallisation) and radiative or non-radiative transition pathways of **PAs** and **BPA**s should be addressed, as well as the reasons for the high fluorescence of **1,4-BPA** and **9,10-BPA** in the solid state.

Firstly, it is important to understand how aggregation changes the fluorescent species in **PAs** and **BPA**s. The relationship between the aggregation state of the solids (*i.e.* colloidal suspension, or neat powder) and the fluorescence quantum yield should provide an insight into the fluorescent species in the solid state. If the aggregation-induced emission of **1,4-BPA** and **9,10-BPA** involves highly delocalised excitons, as reported for anthracenes and typical aromatic hydrocarbons, the fluorescence quantum yields should strongly depend on the conditions of the solids.⁵ However, the fluorescence quantum yields for almost all **PAs** and **BPA**s (except **9-PA**) remain essentially



Fig. 2 Fluorescence spectra of (a) **1,4-BPA** and (b) **9,10-BPA** in THF–water mixtures. Maximum fluorescence intensity of (c) **1,4-BPA** and (d) **9,10-BPA** as a function of water fraction [vol%].



unaffected by the condition of the solid (Table 1). Therefore, the fluorescent excitons of **1,4-BPA** and **9,10-BPA** are likely to be localised on a single, or very few chromophores.

Interestingly, **9-PA** exhibits strong fluorescence ($\Phi_{\text{fl}} \approx 0.71$) only as a neat powder, but not in a suspension. The X-ray crystallographic analysis from a single crystal of **9-PA** revealed a herringbone structure (Fig. S62, ESI[†]), wherein the anthracene moieties adopt a configuration that avoids π -stacking, which is disadvantageous with respect to the formation of strong intermolecular interactions and Frenkel excitons.³⁸ However, given the lack of spectroscopic data for large and high-quality single crystals, which should be necessary to discuss the solid-state fluorescence properties of **9-PA** accurately, we refrained from addressing this question herein.

Accordingly, the diffuse-reflectance spectra of the **PA** and **BPA** derivatives (Fig. S45–S52, ESI[†]) are comparable to their absorption spectra, and peaks associated with delocalised Frenkel excitons, *i.e.* J- or H-aggregates, were not observed.³⁹ Thus, it seems likely that the aggregation does not significantly influence the fluorescent species in **PAs** and **BPAs**. Given that aggregation does not strongly affect radiative pathways, the aggregation effect on the non-radiative transition pathways of **PAs** and **BPAs** should be more important. In order to analyse the effect of aggregation on radiative and non-radiative transition rates, fluorescence lifetime measurements were carried out. However, as shown in Table S1 (ESI[†]), the fluorescence lifetimes measured in colloidal suspensions consisted of two components. Therefore, the lifetimes of bi-exponential decays have to be approximated to a single component to derive the transition rate constants. Sillen *et al.* have reported that the amplitude average lifetime is an effective parameter to derive the transition rate constants of multi-exponential decays caused by the exposure of fluorescent molecules to different environments.⁴⁰ As colloidal suspensions contain at least two separate environments (*i.e.* the aggregation and solution state), the amplitude average lifetime was used to derive the transition rate constants. The derived radiative transition rate constants (k_{r}) remain constant at $\sim 3\text{--}8 \times 10^7 \text{ s}^{-1}$, independent of the sample conditions such as solvent polarity or aggregation. On the other hand, the non-radiative transition rate constants (k_{nr}) for the aggregation state of **PA** ($4\text{--}9 \times 10^7 \text{ s}^{-1}$) and **BPA** ($2\text{--}3 \times 10^8 \text{ s}^{-1}$) are an order of magnitude higher than the corresponding k_{nr} values in solution (**PA**: $3\text{--}9 \times 10^6 \text{ s}^{-1}$; **BPA**: $0.4\text{--}5 \times 10^7 \text{ s}^{-1}$). The exceptions here are **1,4-BPA** and **9,10-BPA**, which exhibit exclusively AIE; their k_{nr} values ($\sim 2\text{--}3 \times 10^7 \text{ s}^{-1}$) suggest that their non-radiative transitions are significantly slower than those of other **BPA** derivatives (see Table S1, ESI[†]).

Thus, it is feasible to conclude that **1,4-BPA** and **9,10-BPA** operate a fluorescence mechanism that avoids the non-radiative transition pathway, which is intrinsic to solid-state **PAs** and **BPAs**. Due to this unique mechanism, aggregation only serves to restrict the intramolecular vibrations, which trigger AIE phenomena.^{1,2}

Relaxation of the excited state in **1,4-BPA** and **9,10-BPA**

In order to extend this method and prepare a variety of simple but bright AIE luminogens based on **1,4-BPA** and **9,10-BPA**, a

concrete design strategy must be established. However, prior to doing this, the intrinsic non-radiative transition pathway of **BPAs** and the mechanism by which **1,4-BPA** and **9,10-BPA** overcome deactivation should be determined.

Usually, non-radiative transition pathways in the solid state arise from energy transfers involving Forster or Dexter mechanisms,^{6,41} excimer or exciplex formation,⁶ the presence of polaronic species,⁴² or self-absorption.⁵ However, the photophysical data obtained in this study were unable to provide consistent and/or direct evidence for a specific mechanism underlying the fluorescence quenching of **PA** and **BPA** derivatives in the solid state. For instance, if we assume that the non-radiative pathway occurs by self-absorption, a colloidal suspension should experience a stronger bathochromic shift of its fluorescence relative to those of diluted and non-polar solutions. Even though **1,5-BPA** and **1,8-BPA** follow this paradigm, suspensions of **2,6-BPA** exhibit a more hypsochromically shifted fluorescence emission than those of the solutions (Fig. S35–S44, ESI[†]). Accordingly, it is hardly debatable that further studies, focusing on a detailed analysis of the photophysical processes in the solid state, are required.

Nevertheless, the efficiency of the non-radiative transitions can be elucidated from X-ray crystallographic analyses and from the photophysical properties of dilute solutions. Fig. 3 shows the molecular structures of **1,4-BPA**, **1,5-BPA**, and **9,10-BPA**, obtained from single crystal X-ray diffraction studies. It is widely accepted that undesirable interchromophoric interactions can be suppressed by separating chromophores spatially. Hence, a commonly employed design principle for the enhancement of solid-state emission is based on the use of bulky substituents and multiple aryl–aryl bonds.^{6,7} However, **BPAs** lack such substituents and bonds, and the anthracene chromophores are packed in such close proximity that Dexter-type energy transfers should be possible (*i.e.* interchromophoric distance $\leq 10 \text{ \AA}$).⁴¹ Even though molecules of **1,4-BPA** are in close proximity in the crystal structure, their orthogonal



Fig. 3 Molecular packing in the unit cell of (a) **1,4-BPA**, (b) **9,10-BPA**, and (c) **1,5-BPA**. The number values refer to the closest distances (in \AA) between neighbouring chromophores.



packing may minimise interchromophoric interactions, while simultaneously enhancing the fluorescence brightness (Fig. 3a). As the unit cells of the **9,10-BPA** crystals lack a molecule at the body-centred position, **9,10-BPA** molecules are packed in a face-to-face manner and arranged more closely compared to **1,4-BPA** and **1,5-BPA** (Fig. 3b). Thus, when only the electronic interaction with neighbouring chromophores is considered, **9,10-BPA** is more likely to undergo fluorescence quenching than other **BPA**s. However, close packing is, at the same time, favourable to suppress internal conversion. Recent theoretical studies⁴³ demonstrated that a close packing in the solid state effectively restricts a low-frequency vibrational mode, which is strongly coupled with the reorganization energy that plays a pivotal role in fast internal conversions. This hypothesis can be applied to the AIE behaviour of **9,10-BPA**: provided that the internal conversion rate of **9,10-BPA** is, as for 9,10-diphenylanthracene,⁴⁴ governed by a low-frequency mode such as a C–N rotation, it is reasonable to assume that the close packing of **9,10-BPA** leads to a slow non-radiative transition. To quantitatively analyse the effect of the close packing, we will conduct experimental and theoretical investigations using model compounds.

It is noteworthy that the shape of the fluorescence spectra (Fig. S53–S60, ESI†) for several **PA**s and **BPA**s significantly depends on the aggregation condition of the solid (*i.e.* colloidal suspension, 1 mM in NaBr, or neat powder), which is indicative of their mechanofluorochromic properties. However, this spectral change is less pronounced for AIE-active **1,4-BPA** and **9,10-BPA**. To impart these AIE-active dyes with mechanofluorochromic properties, we designed **1,4-BPA** and **9,10-BPA** analogues with a bulkier dialkylamine group.⁴⁵

The photophysical properties of dilute solutions, especially the Stokes shifts, are also relevant to the non-radiative transition rates in the solid state.⁴⁷ Large Stokes shifts imply a marginal overlap between the absorption and fluorescence spectra, which prevents energy transfer⁴¹ and self-absorption. Fig. 4 shows a plot of the Stokes shifts of some **PA** and **BPA** derivatives in toluene, THF, and DMF as a function of solvent polarity parameter (f_L) of each solvent.⁴⁸ Since the Stokes shifts and other spectroscopic features of **PA** are similar to those of 1-, 2-, and 9-aminoanthracenes⁴⁹ and their

dimethylamine analogs,^{21,50} absorption and fluorescence spectra were assigned accordingly. While $S_0 \rightarrow S_1$ absorptions (Fig. S27 and S33, ESI†) of **1-PA** and **9-PA** stem from symmetry-allowed mixing between the 1L_a state of anthracene⁵¹ and the 1CT (singlet charge transfer) state, a more red-shifted $S_0 \rightarrow S_1$ absorption (Fig. S31, ESI†) is observed for **2-PA**, which results from the mixing between the optically forbidden 1L_b state of anthracene and its 1CT state.⁵¹ Thus, the resultant Stokes shift of **2-PA** is smaller than **1-PA** despite small steric hindrance against C–N rotation. As the character of the $S_0 \rightarrow S_1$ absorption/fluorescence is very different for each **PA**, the effect of introducing a second piperidine group should be evaluated by comparison with the corresponding **PA**. As shown in Fig. 4, the Stokes shifts of **1,5-BPA**, **1,8-BPA**, and **2,6-BPA** are smaller than those of the corresponding monosubstituted analogues, *i.e.* **1-PA** and **2-PA**. This decreased Stokes shift can be ascribed to the symmetric excited-state dipole moment of **1,5-BPA**, **1,8-BPA**, and **2,6-BPA**, as their Stokes shifts are less dependent on solvent polarity than those of the **PA**s. Conversely, **1,4-BPA** and **9,10-BPA** exhibited Stokes shifts, which are significantly larger than those of **1-PA** and **9-PA**, even though the Stokes shifts of these **BPA**s are equally less dependent on solvent polarity than corresponding **PA**s. In dilute solution, the Stokes shift of a fluorophore is determined by dielectric interactions with solvent molecules (outer-sphere reorganisation) and the deformation of solute molecules induced by excited-state energy relaxation (inner-sphere reorganisation). When considering small degrees of solvatochromism for **BPA**s into account, prominent Stokes shifts of **1,4-BPA** and **9,10-BPA** should arise from their large inner-sphere reorganization energy. It can therefore be concluded that only the second piperidyl groups at the *para*-position, in *e.g.* **1,4-BPA** and **9,10-BPA**, contribute significantly to the excited-state energy relaxation and hence induce a larger Stokes shift.

In order to gain a better understanding of the origin of the large Stokes shifts in **1,4-BPA**, we performed time-dependent density functional theory (TD-DFT) calculations for **1,4-BPA** and **1,5-BPA** at the ω B97X-D/6-311G(d,p) level of theory (see ESI†) and the results are summarised in Fig. 5. The calculated HOMOs and LUMOs for both **1,4-BPA** and **1,5-BPA** are derived from the anthracene $^1b_{2g}$ and $^1b_{3u}$ orbital, respectively.⁵²

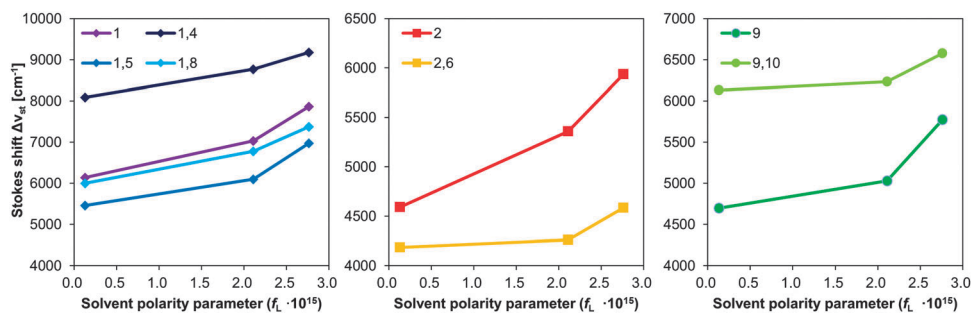


Fig. 4 Stokes shifts of **PA** and **BPA** derivatives plotted against the solvent polarity parameter (f_L).²⁴ The solvent polarity parameter is obtained from the following equation: $f_L = 2\Delta f/hc$, $\Delta f = (\epsilon - 1)/(2\epsilon + 1) - (n^2 - 1)/(2n^2 + 1)$, where c refers to the speed of light [cm s⁻¹], h to the Planck constant [erg s], n to the refractive index,⁴⁶ and ϵ to the dielectric constant of each solvent.⁴⁶





Fig. 5 TD-DFT calculations for the energy levels of the HOMO and LUMO of the equilibrium structures for the ground (S_0) and excited state (S_1), as well as the corresponding wavelengths for the S_0 – S_1 transitions (top): Kohn–Sham orbitals for the HOMO of each structure. All calculations were carried out at the ω B97X-D/6-311G(d,p) level of theory.

Thus, the HOMO–LUMO transitions in these **BPA**s, which lead to the first singlet excited states S_1 , are based on the S_0 – 1L_a transition (π – π^* character) of anthracene⁵³ (Tables S7 and S8, ESI†).

The Stokes shifts for the S_0 – S_1 transition in **1,4-BPA** are significantly larger than for that in **1,5-BPA**. The distribution of the HOMO does not show any significant change upon relaxation of **1,5-BPA**, whereas the HOMO of **1,4-BPA** clearly converges on the substituted benzene ring of the anthracene moiety after the structural relaxation. It can thus be concluded that the large Stokes shift of **1,4-BPA** originates from the destabilisation of the HOMO, which is caused by the strong electronic repulsion between two piperidyl groups in close spatial proximity.

For the evaluation of the structural change induced by the inner-sphere reorganisation, we calculated the change of the dihedral angle between the anthracene plane (defined by atoms 4, 12, and 14; Tables S3–S6, ESI†) and the piperidyl group (defined by atoms 14, 18, and 20; Tables S3–S6, ESI†). For **1,5-BPA**, dihedral angles of 51° (S_0) and 44° (S_1) were obtained, while angles of 52° (S_0) and 39° (S_1) were obtained for **1,4-BPA**. Twisted intramolecular charge-transfer (TICT) states often display excited-state species with large Stokes shifts,⁵⁴ and their inner-sphere reorganisation energy is comparable to that of **1,4-BPA**. However, the inner-sphere reorganisation of TICT dyes demands a significantly larger change of the torsion angle between the donor and acceptor.^{55,56} Notably, the change of the torsion angle from S_0 geometry to S_1 geometry is only moderate for **1,4-BPA** (13°), which suggests that **1,4-BPA** can, in contrast to TICT dyes, complete the inner-sphere reorganisation even under restricted mobility in the solid state. In fact, the restricted environment in the solid state prohibits the formation of the TICT state, which frequently results in AIE,^{12,57} albeit concomitant with decreased Stokes shifts due to the limited inner-sphere reorganisation in the excited state. This trade-off relationship in TICT dyes requires elaborate and

prudent design strategies.³⁴ Given that electronic repulsion is able to induce large Stokes shifts with a minimal structural relaxation, twisted donors in spatial proximity such as in **1,4-BPA** represent a simple, yet viable option to simultaneously obtain effective AIE and large Stokes shifts.

Although we performed TD-DFT calculations for **9,10-BPA** as well, the results were so complicated that we were unable to interpret the underlying photophysical processes comprehensively. Considering its crystal structure, the mechanism for the solid-state emission and AIE in **9,10-BPA** may be different from that in **1,4-BPA**. Future studies will investigate the photophysical properties of **9,10-BPA** by spectroscopic measurements, the synthesis of model compounds, and computational calculations.

Conclusion

In this study, we synthesised various regioisomers of piperidylanthracene (**PA**) and bis(piperidyl)anthracene (**BPA**), and evaluated their photophysical properties. The result revealed that only **1,4-BPA** and **9,10-BPA** exhibit AIE, and that their fluorescence colour can be tuned by changing the relative position of the piperidyl groups.

Results from diffuse-reflectance spectroscopy and fluorescence lifetime measurements demonstrated that **PA** and **BPA** intrinsically possess an efficient non-radiative transition pathway in the solid state, and that **1,4-BPA** and **9,10-BPA** may overcome this pathway. While the molecular packing of **1,4-BPA** in the crystalline state exhibited weak interchromophoric interactions, that of **9,10-BPA** is indicative of the importance of suppression of intramolecular vibrations. Spectroscopic measurements in dilute solutions indicated that the Stokes shifts of **1,4-BPA** and **9,10-BPA** are considerably larger than those of their corresponding **PA** analogues. Accordingly, the spectroscopic characteristics of **1,4-BPA** and **9,10-BPA** are suitable for the suppression of self-absorption and energy transfer in the solid state. TD-DFT calculations corroborated that the large Stokes shift of **1,4-BPA** originates from a drastic destabilisation of the HOMO, induced by a minimal structural relaxation in the excited state.

Nevertheless, several important points, such as *e.g.* the intrinsic quenching process of **PAs** and **BPA**s in the solid state, the mechanism, by which **9,10-BPA** overcomes its quenching path with a close face-to-face packing, as well as the relaxation process of the excited state of **9,10-BPA** still remain to be revealed. Investigations in these directions are currently in progress in our laboratory, and we also aim to apply the excellent photophysical properties of **1,4-BPA** and **9,10-BPA** in various functional molecules modifying dialkylamine substituents.

Experimental section

Instruments

^1H NMR and ^{13}C NMR spectra were recorded on a 400 MHz JEOL LMN-EX400 or a 300 MHz BRUKER DPX300 spectrometer



using tetramethylsilane (TMS) as the internal standard. FT-IR spectra were recorded on a JASCO FT-IR 469 plus spectrometer. Melting points were recorded on a Yanaco micro melting point apparatus MP-500P. Thermogravimetric analyses were carried out on a TG/DTA 6200 SII Seiko Instruments Inc. EXSTAR 6000 thermal analyzer at a heating rate of 5 °C min⁻¹ under a nitrogen atmosphere. MS spectra (FAB⁺) were obtained using a JEOL JMS700 mass spectrometer.

All photophysical measurements performed in solution were carried out using dilute solutions with optical densities (ODs) around 0.1 at the maximum absorption wavelength in 1 cm path length quartz cells at room temperature (298 K). In addition, toluene, THF, and DMF solutions of all samples were deaerated by bubbling with argon gas for 15 min prior to the measurements. For photophysical measurements of the aggregate state, colloidal suspensions (THF:H₂O = 1:9; v/v) with a concentration of 1.0 × 10⁻⁴ M were prepared, and aggregate formation was confirmed by the severe broadening of the UV-vis spectra. All photophysical measurements of these colloidal suspensions were conducted immediately after the aggregate formation in order to avoid experimental errors due to surface oxidation. UV-vis spectra were recorded on a JASCO V-670 UV-vis spectrophotometer. Fluorescence spectra were recorded on a JASCO FP-6500 spectrofluorometer. The wavelengths obtained by the fluorescence spectrometer were converted to the wavenumber using the equation $I(\tilde{\nu}) = \lambda^2 I(\lambda)$.⁵⁸ Absolute quantum yields were measured using a Hamamatsu Photonics Quantaurus QY apparatus. Fluorescence lifetime values were obtained from the most intense peaks using a Hamamatsu Photonics OB 920 fluorescence lifetime spectrometer equipped with an LED lamp (343 nm).

Diffuse-reflectance spectra

Diffuse-reflectance spectra were recorded on a JASCO FP-6500 spectrofluorometer equipped with an integration sphere detector. Thus, the experimental error arising from the fluorescence, typically encountered with diffuse-reflectance spectrometers producing polychromatic outgoing light, was avoided. Samples and references were charged in a JASCO powder sample cell to obtain a sufficiently thick powder layer. Just before each measurement, the synchronous spectrum of the NaBr powder was measured as a reflectance spectrum of a standard reference $r_{\text{standard}}(\lambda)$. Then, a reflectance spectrum of each sample $r_{\text{sample}}(\lambda)$ was obtained following the same procedure. In addition to neat powder samples, the spectra of samples adsorbed on the NaBr powder were measured at a concentration of 1.0 × 10⁻³ M. The obtained reflectance spectra $r_{\text{sample}}(\lambda)$ and $r_{\text{standard}}(\lambda)$ were converted to Kubelka–Munk functions $f(r_{\infty})$ using the following equation:⁵⁹

$$f(r_{\infty}) = \frac{(1 - r_{\infty}(\lambda))^2}{2r_{\infty}(\lambda)}, \quad \text{where } r_{\infty} = \frac{r_{\text{sample}}(\lambda)}{r_{\text{standard}}(\lambda)}$$

All diffuse-reflectance spectra are displayed as plots of the Kubelka–Munk functions, *i.e.* $f(r_{\infty})$ as a function of the wavelength, λ .

DLS measurement

All of DLS measurements were performed on a Wyatt DynaPro NanoStar. Autocorrelation functions attained from DLS measurement (Fig. S18, ESI[†]) clearly indicated the formation of aggregates in THF:H₂O = 1:9 (v/v), and the autocorrelation functions of all sample suspension are smooth and continuous, exhibiting exponential decay from their maxima to a value of 1. Blank solutions (THF:H₂O = 1:9; v/v) did not show any significant autocorrelation.

The particle size distribution (Fig. S26–S29, ESI[†]) was calculated by a regularisation analysis⁶⁰ of the autocorrelation functions. Bimodal distributions were observed for **2-PA** and **1,8-BPA**, while other **PA** and **BPA** regioisomers exhibited unimodal distributions. Percentage polydispersities (*i.e.* the width of each peak normalised to the mean size of the peak) of all peaks were smaller than 20%, and hydrodynamic radii R_h are included in the particle size distributions.

Materials

Unless otherwise noted, all reagents and chemicals were used as received without further purification. 1-Bromoanthracene, 2-bromoanthracene, 9-bromoanthracene, 1,4-diaminoanthraquinone, 1,5-dibromoanthracene, 1,8-dibromoanthracene, 2,6-dibromoanthracene, 9,10-dibromoanthracene, sodium *tert*-butoxide and (±)-2,2'-bis(diphenylphosphino)-1,1'-binaphthyl (BINAP) were obtained from TCI (Tokyo, Japan). The PEPPSI[™]-IPr catalyst was purchased from Sigma-Aldrich Japan (Tokyo, Japan). Palladium(II)diacetate and dehydrated 1,4-dioxane were purchased from Wako Pure Chem (Tokyo, Japan). Dehydrated THF and copper(II)bromide were obtained from Kanto Chem (Tokyo, Japan). Spectrograde toluene, THF, DMF, magnesium sulphate and sodium carbonate were purchased from Nacalai Tasque (Kyoto, Japan). Pure water used for photophysical measurements was prepared using a Milli-Q[®] water purification system (Merck Millipore, United States).

Synthesis and characterization of PAs and BPAs

¹H NMR spectra and ¹³C NMR spectra are summarized in the ESI[†].

1-Piperidylanthracene (1-PA). A solution of sodium *tert*-butoxide (0.56 g, 5.8 mmol), the PEPPSI[™]-IPr catalyst (48 mg, 0.07 mmol), and 1-bromoanthracene (0.30 g, 1.17 mmol) in deaerated and dehydrated 1,4-dioxane (30 mL) was stirred for 15 min under argon at room temperature, before piperidine (0.35 mL, 3.50 mmol) was added. The resulting mixture was stirred for 24 h at 100 °C. After cooling to room temperature, the reaction was quenched by the addition of water, before the mixture was filtered. Subsequently, the filtrate was extracted with chloroform, and the organic layer was dried over magnesium sulfate. The solvent was removed *in vacuo*, and the residue was purified by column chromatography on silica gel using chloroform:hexane = 2:3 as the eluent to yield crude **1-PA**. The crude product thus obtained was further purified by recrystallisation from hexane to afford **1-PA** (0.30 g) as a pale yellow solid. Yield: 98%; mp: 83.4–84.1 °C; ¹H NMR (400 MHz, CDCl₃):



δ 8.74 (s, 1H), 8.39 (s, 1H), 8.04 (dd, $^4J = 2.9$ Hz, $^3J = 6.0$ Hz, 1H), 7.97 (dd, $^4J = 2.9$ Hz, $^3J = 6.0$ Hz, 1H), 7.68 (d, $^3J = 8.4$ Hz, 1H), 7.44 (dd, $^4J = 2.9$ Hz, $^3J = 6.4$ Hz, 2H), 7.39–7.35 (m, 1H), 6.99 (d, $^3J = 7.1$ Hz, 1H), 3.13 (br, 4H), 1.92 (m, 4H), 1.72 (m, 2H) ppm (Fig. S1, ESI[†]); ^{13}C NMR (100 MHz, CDCl_3): δ 151.0 (Ar), 133.0 (Ar), 131.5 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 131.2 (Ar), 128.8 (Ar), 127.9 (Ar), 127.8 (Ar), 126.5 (Ar), 125.4 (Ar), 125.3 (Ar), 125.0 (Ar), 123.1 (Ar), 122.6 (Ar), 113.0 (Ar), 54.6 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 26.7 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 24.7 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$) ppm (Fig. S2, ESI[†]); FT-IR (KBr): 2933 ($-\text{CH}_2-$ stretch), 1615 (Ar ring stretch), 1537 (Ar ring stretch), 1401 (Ar–N stretch), 735 (Ar–H) cm^{-1} ; HRMS (FAB⁺) exact mass calculated for $[\text{M}]^+$ ($\text{C}_{19}\text{H}_{19}\text{N}$) requires m/z 261.1517, found m/z 261.1512.

1,4-Bis(piperidyl)anthracene (1,4-BPA). 1,4-Dibromoanthracene, the precursor for 1,4-bis(piperidyl)anthracene, was prepared from 1,4-diaminoanthraquinone according to literature procedures.^{61,62} The same procedure as described for the preparation of **1-PA** was employed for the synthesis of **1,4-BPA** using sodium *tert*-butoxide (0.47 g, 4.9 mmol), the PEPPSI[™]-IPr catalyst (40 mg, 0.06 mmol), 1,4-bromoanthracene (0.33 g, 0.97 mmol), 30 mL of 1,4-dioxane, and piperidine (0.29 mL, 2.9 mmol). Purification by column chromatography on silica gel ((1) chloroform : hexane = 1 : 4; (2) chloroform : hexane = 1 : 1), followed by recrystallisation of the crude product from hexane to afford **1,4-BPA** (0.28 g) as a yellow solid. Yield: 83%; mp: 225.3–228.7 °C (decomposition onset at $T = 200$ °C); ^1H NMR (300 MHz, CDCl_3): δ 8.78 (s, 2H), 8.05 (dd, $^4J = 3.2$ Hz, $^3J = 6.4$ Hz, 2H), 7.47 (dd, $^4J = 3.2$ Hz, $^3J = 6.4$ Hz, 2H), 6.96 (s, 2H), 3.27 (br, 8H), 1.92 (m, 8H), 1.72 (br, 4H) ppm (Fig. S3, ESI[†]); ^{13}C NMR (75 MHz, CDCl_3): δ 146.8 (Ar), 131.2 (Ar), 129.2 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 128.7 (Ar), 125.2 (Ar), 122.9 (Ar), 113.3 (Ar), 55.0 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 26.9 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 24.9 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$) ppm (Fig. S4, ESI[†]); FT-IR (KBr): 2935 ($-\text{CH}_2-$ stretch), 1610 (Ar ring stretch), 1335 (Ar–N stretch), 745 (Ar–H) cm^{-1} ; HRMS (FAB⁺) exact mass calculated for $[\text{M}]^+$ ($\text{C}_{24}\text{H}_{28}\text{N}_2$) requires m/z 344.2252, found m/z 344.2259.

1,5-Bis(piperidyl)anthracene (1,5-BPA). Compound **1,5-BPA** was prepared as described for **1-PA** using sodium *tert*-butoxide (0.72 g, 7.5 mmol), the PEPPSI[™]-IPr catalyst (61 mg, 0.09 mmol), 1,5-dibromoanthracene (0.50 g, 1.5 mmol), 40 mL of 1,4-dioxane, and piperidine (0.44 mL, 4.5 mmol) to afford crude **1,5-BPA**. The crude product was purified by column chromatography on silica (chloroform : hexane = 1 : 1), followed by recrystallisation from hexane to afford **1,5-BPA** (0.47 g) as a yellow solid. Yield: 91%; mp: 248.5–249.4 °C; ^1H NMR (300 MHz, CDCl_3): δ 8.68 (s, 2H), 7.71 (dd, $^4J = 0.9$ Hz, $^3J = 8.4$ Hz, 2H), 7.36 (dd, $^3J = 7.2$ Hz, $^3J = 8.4$ Hz, 2H), 6.97 (dd, $^4J = 0.9$ Hz, $^3J = 7.2$ Hz, 2H), 3.13 (br, 8H), 1.93–1.88 (m, 8H), 1.71 (br, 4H) ppm (Fig. S5, ESI[†]); ^{13}C NMR (100 MHz, CDCl_3): δ 150.9 (Ar), 133.7 (Ar), 127.9 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 125.2 (Ar), 123.9 (Ar), 123.1 (Ar), 113.0 (Ar), 54.7 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 26.8 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 24.9 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$) ppm (Fig. S6, ESI[†]); FT-IR (KBr): 2934 ($-\text{CH}_2-$ stretch), 1614 (Ar ring stretch), 1540 (Ar ring stretch), 1348 (Ar–N stretch), 733 (Ar–H) cm^{-1} ; HRMS (FAB⁺) exact mass calculated for $[\text{M}]^+$ ($\text{C}_{24}\text{H}_{28}\text{N}_2$) requires m/z 344.2252, found m/z 344.2259.

1,8-Bis(piperidyl)anthracene (1,8-BPA). Compound **1,8-BPA** was prepared as described for **1-PA** using sodium *tert*-butoxide (0.72 g, 7.5 mmol), the PEPPSI[™]-IPr catalyst (61 mg, 0.09 mmol), 1,8-dibromoanthracene (0.50 g, 1.5 mmol), 40 mL of 1,4-dioxane, and piperidine (0.44 mL, 4.5 mmol). Purification by column chromatography on silica gel (chloroform : hexane = 2 : 3), followed by recrystallisation from hexane to yield **1,8-BPA** (0.41 g) as a yellow solid. Yield: 80%; mp: 248.1–249.2 °C (the initially yellow solid turned into a white crystalline solid at 247.5 °C); ^1H NMR (300 MHz, CDCl_3): δ 9.17 (s, 1H), 8.34 (s, 1H), 7.65 (d, $^3J = 8.4$ Hz, 2H), 7.36 (dd, $^3J = 7.1$ Hz, $^3J = 8.4$ Hz, 2H), 6.97 (d, $^3J = 7.1$ Hz, 2H), 3.14 (br, 8H), 1.95–1.93 (m, 8H), 1.72 (br, 4H) ppm (Fig. S7, ESI[†]); ^{13}C NMR (100 MHz, CDCl_3): δ 151.6 (Ar), 133.0 (Ar), 127.5 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 126.9 (Ar), 125.5 (Ar), 123.0 (Ar), 119.1 (Ar), 113.1 (Ar), 54.9 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 27.1 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 25.0 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$) ppm (Fig. S8, ESI[†]); FT-IR (KBr): 2938 ($-\text{CH}_2-$ stretch), 1615 (Ar ring stretch), 1562, 1448, 1332 (Ar–N stretch), 1244, 896, 743 (Ar–H) cm^{-1} ; HRMS (FAB⁺) exact mass calculated for $[\text{M}]^+$ ($\text{C}_{24}\text{H}_{28}\text{N}_2$) requires m/z 344.2252, found m/z 344.2259. These characterization data are consistent with those reported in the literature.^{28a}

2-Piperidylanthracene (2-PA). Compound **2-PA** was prepared as described for **1-PA** using sodium *tert*-butoxide (0.93 g, 9.7 mmol), the PEPPSI[™]-IPr catalyst (79 mg, 0.12 mmol), 2-bromoanthracene (0.50 g, 1.9 mmol), 40 mL of 1,4-dioxane, and piperidine (0.58 mL, 5.8 mmol). Purification by column chromatography on silica gel (chloroform : hexane = 1 : 3), followed by recrystallisation from hexane to afford **2-PA** (0.32 g) as a pale yellow solid. Yield: 63%; mp: 168.1–169.1 °C (lit. 162–164 °C (ref. 63)); ^1H NMR (400 MHz, CDCl_3): δ 8.27 (s, 1H), 8.19 (s, 1H), 7.92 (dd, $^3J = 7.8$ Hz, $^3J = 7.8$ Hz, 2H), 7.86 (d, $^3J = 9.3$ Hz, 1H), 7.42–7.32 (m, 3H), 7.17 (d, $^4J = 2.0$ Hz, 1H), 7.39–7.35 (m, 1H), 6.99 (d, $^3J = 7.1$ Hz, 1H), 3.30 (t, $^3J = 5.4$ Hz, 4H), 1.82–1.76 (m, 4H), 1.67–1.63 (m, 2H) ppm (Fig. S9, ESI[†]); ^{13}C NMR (100 MHz, CDCl_3): δ 149.5 (Ar), 133.3 (Ar), 132.4 (Ar), 130.3 (Ar), 129.0 (Ar), 128.3 (Ar), 128.1 (Ar), 127.7 (Ar), 125.9 (Ar), 125.3 (Ar), 124.1 (Ar), 123.9 (Ar), 121.8 (Ar), 108.4 (Ar), 51.0 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 26.1 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$), 24.5 ($\text{ArN}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)$) ppm (Fig. S10, ESI[†]); FT-IR (KBr): 2932 ($-\text{CH}_2-$ stretch), 1625 (Ar ring stretch), 1483 (Ar ring stretch), 1344 (Ar–N stretch), 741 (Ar–H) cm^{-1} ; HRMS (FAB⁺) exact mass calculated for $[\text{M}]^+$ ($\text{C}_{19}\text{H}_{19}\text{N}$) requires m/z 261.1517, found m/z 261.1523.

2,6-Bis(piperidyl)anthracene (2,6-BPA). Compound **2,6-BPA** was prepared as described for **1-PA** using sodium *tert*-butoxide (0.72 g, 7.5 mmol), the PEPPSI[™]-IPr catalyst (61 mg, 0.09 mmol), 2,6-dibromoanthracene (0.50 g, 1.5 mmol), 40 mL of 1,4-dioxane, and piperidine (0.44 mL, 4.5 mmol). Purification by column chromatography on silica gel (chloroform), followed by consecutive recrystallisations from hexane–chloroform and hexane to afford **2,6-BPA** (0.28 g) as a yellow solid. Yield: 54%; mp: 305.2–306.0 °C (the sample quickly decomposed after melting); ^1H NMR (300 MHz, CDCl_3): δ 8.07 (s, 2H), 7.79 (d, $^3J = 9.2$ Hz, 2H), 7.28 (dd, $^4J = 2.3$ Hz, $^3J = 9.2$ Hz, 2H), 7.14 (d, $^4J = 2.3$ Hz, 2H), 3.25 (t, $^3J = 5.5$ Hz, 8H), 1.82–1.75 (m, 8H), 1.66–1.61 (m, 4H) ppm (Fig. S11, ESI[†]); ^{13}C NMR (100 MHz, CDCl_3): δ 148.7



(Ar), 131.7 (Ar), 128.7 (Ar), 128.5 (Ar), 123.8 (Ar), 121.8 (Ar), 109.2 (Ar), 51.3 (ArN(CH₂CH₂CH₂CH₂CH₂)), 26.1 (ArN(CH₂CH₂CH₂CH₂CH₂)), 24.5 (ArN(CH₂CH₂CH₂CH₂CH₂)) ppm (Fig. S12, ESI[†]); FT-IR (KBr): 2930 (–CH₂– stretch), 1626 (Ar ring stretch), 1382 (Ar–N stretch) cm^{–1}; HRMS (FAB⁺) exact mass calculated for [M]⁺ (C₂₄H₂₈N₂) requires *m/z* 344.2252, found *m/z* 344.2259.

9-Piperidylanthracene (9-PA). Compound **9-PA** was prepared as described for **1-PA** using sodium *tert*-butoxide (0.22 g, 2.3 mmol), the PEPPSI[™]-IPr catalyst (79 mg, 0.06 mmol), 9-bromoanthracene (0.30 g, 1.2 mmol), 20 mL of 1,4-dioxane, and piperidine (1.38 mL, 14 mmol). Purification by column chromatography on silica gel (chloroform : hexane = 1 : 3), followed by recrystallisation from hexane to afford **9-PA** (0.088 g) as a pale yellow solid. Yield: 29%; mp: 145.2–146.0 °C; ¹H NMR (300 MHz, CDCl₃): δ 8.52–8.48 (m, 2H), 8.27 (s, 1H), 8.01–7.96 (m, 2H), 7.48–7.40 (m, 4H), 3.47 (t, ³J = 5.0 Hz, 4H), 1.92–1.77 (m, 6H) ppm (Fig. S13, ESI[†]); ¹³C NMR (100 MHz, CDCl₃): δ 145.8 (Ar), 132.6 (Ar), 130.5 (Ar), 128.8 (Ar), 125.1 (Ar), 125.0 (Ar), 124.7 (Ar), 124.3 (Ar), 53.1 (ArN(CH₂CH₂CH₂CH₂CH₂)), 27.5 (ArN(CH₂CH₂CH₂CH₂CH₂)), 24.9 (ArN(CH₂CH₂CH₂CH₂CH₂)) ppm (Fig. S14, ESI[†]); FT-IR (KBr): 2939 (–CH₂– stretch), 1623 (Ar ring stretch), 1491 (Ar ring stretch), 735 (Ar–H) cm^{–1}; HRMS (FAB⁺) exact mass calculated for [M]⁺ (C₁₉H₁₉N) requires *m/z* 261.1517, found *m/z* 261.1524.

9,10-Bis(piperidyl)anthracene (9,10-BPA). A solution of sodium *tert*-butoxide (0.40 g, 4.2 mmol), palladium acetate (33 mg, 0.15 mmol), 9,10-dibromoanthracene (0.5 g, 1.5 mmol), and BINAP (0.29 g, 0.44 mmol) in toluene (15 mL) was stirred for 15 min under argon at 100 °C. Then, piperidine (0.59 mL, 6.0 mmol) was added. The resulting mixture was stirred at 100 °C for 72 h, before the reaction was quenched with water. The mixture was filtered and the filtrate was extracted with chloroform, before the organic layer was dried over magnesium sulfate. After filtration, the solvent was removed *in vacuo*, and the residue was purified by passage through a short silica gel plug followed by high-performance column chromatography (chloroform) to afford crude **9,10-BPA**. The crude product thus obtained was further purified by recrystallisation (hexane : ethyl acetate = 2 : 1) to afford **9,10-BPA** (310 mg) as a yellow solid. Yield: 60%; ¹H NMR (300 MHz, CDCl₃): δ 8.50 (dd, ⁴J = 3.4 Hz, ³J = 6.8 Hz, 4H), 7.43 (dd, ⁴J = 3.3 Hz, ³J = 6.8 Hz, 4H), 3.54 (t, ³J = 4.5 Hz, 8H), 1.86–1.79 (m, 12H) ppm (Fig. S15, ESI[†]); ¹³C NMR (75 MHz, CDCl₃): δ 143.8 (ArN(CH₂CH₂CH₂)), 131.6 (Ar), 125.8 (Ar), 124.6 (Ar), 53.2 (ArN(CH₂CH₂CH₂CH₂CH₂)), 27.8 (ArN(CH₂CH₂CH₂CH₂CH₂)), 25.1 (ArN(CH₂CH₂CH₂CH₂CH₂)) ppm (Fig. S16, ESI[†]); FT-IR (KBr): 2936 (–CH₂– stretch), 2847 (–CH₂– stretch), 1619 (Ar ring stretch), 1401 (Ar–N stretch), 771 (Ar–H) cm^{–1}; HRMS (FAB⁺) exact mass calculated for [M]⁺ (C₂₄H₂₈N₂) requires *m/z* 344.2252, found *m/z* 344.2262. The melting point could not be determined due to decomposition at 280 °C.

References

- 1 Y. Hong, J. W. Y. Lam and B. Z. Tang, *Chem. Soc. Rev.*, 2011, **40**, 5361–5388.
- 2 Y. Hong, J. W. Y. Lam and B. Z. Tang, *Chem. Commun.*, 2009, 4332–4353.
- 3 K. Suzuki, A. Kobayashi, S. Kaneko, K. Takehira, T. Yoshihara, H. Ishida, Y. Shiina, S. Oishi and S. Tobita, *Phys. Chem. Chem. Phys.*, 2009, **11**, 9850–9860.
- 4 (a) L. Porres, A. Holland, L. O. Palsson, A. P. Monkman, C. Kemp and A. Beeby, *J. Fluoresc.*, 2006, **16**, 267–272; (b) Y. Kawamura, H. Sasabe and C. Adachi, *Jpn. J. Appl. Phys.*, 2004, **43**, 7729–7730.
- 5 R. Katoh, K. Suzuki, A. Furube, M. Kotani and K. Tokumaru, *J. Phys. Chem. C*, 2009, **113**, 2961–2965.
- 6 M. Shimizu and T. Hiyama, *Chem. – Asian J.*, 2010, **5**, 1516–1531.
- 7 S. P. Anthony, *ChemPlusChem*, 2012, **77**, 518–531.
- 8 (a) W. Z. Yuan, P. Lu, S. Chen, J. W. Y. Lam, Z. Wang, Y. Liu, H. S. Kwok, Y. Ma and B. Z. Tang, *Adv. Mater.*, 2010, **22**, 2159–2163; (b) J. Huang, N. Sun, P. Chen, R. Tang, Q. Li, D. Ma and Z. Li, *Chem. Commun.*, 2014, **50**, 2136–2138.
- 9 J. Mei, J. Wang, J. Z. Sun, H. Zhao, W. Yuan, C. Deng, S. Chen, H. H. Y. Sung, P. Lu, A. Qin, H. S. Kwok, Y. Ma, I. D. Williams and B. Z. Tang, *Chem. Sci.*, 2012, **3**, 549–558.
- 10 (a) H. Wang, Y. Huang, X. Zhao, W. Gong, Y. Wang and Y. Cheng, *Chem. Commun.*, 2014, **50**, 15075–15078; (b) S. Y. Yoon, J. W. Chung, J. Gierschner, K. S. Kim, M. G. Choi, D. Kim and S. Y. Park, *J. Am. Chem. Soc.*, 2010, **132**, 13675–13683; (c) N. B. Shustova, T. C. Ong, A. F. Cozzolino, V. K. Michaelis, R. G. Griffin and M. Dinca, *J. Am. Chem. Soc.*, 2012, **134**, 15061–15070; (d) A. Battisti, F. Tantussi, F. Fuso, M. Allegrini, G. Ruggeri and A. Pucci, *Macromol. Rapid Commun.*, 2014, **215**, 499–506; (e) X. Yao, X. Ma and H. Yuan, *J. Mater. Chem. C*, 2014, **2**, 5155–5160.
- 11 (a) D. Ding, K. Li, B. Liu and B. Z. Tang, *Acc. Chem. Res.*, 2013, **46**, 2441–2453; (b) X. Zhang, X. Zhang, L. Tao, Z. Chi, J. Xu and Y. Wei, *J. Mater. Chem. B*, 2014, **2**, 4398–4414.
- 12 Y. Dong, J. W. Y. Lam, A. Qin, J. Sun, J. Liu, Z. Li, S. Zhang, J. Sun, H. S. Kwok and B. Z. Tang, *Appl. Phys. Lett.*, 2007, **91**, 011111.
- 13 (a) L. Zhu and Y. Zhao, *J. Mater. Chem. C*, 2013, **1**, 1059–1065; (b) R. Wei, Y. He, X. Wang and P. Keller, *Macromol. Rapid Commun.*, 2014, **35**, 1571–1577.
- 14 (a) X. Zhang, Z. Chi, H. Li, B. Xu, X. Li, S. Liu, Y. Zhang and J. Xu, *J. Mater. Chem.*, 2011, **21**, 1788–1796; (b) X. Zhang, Z. Chi, J. Zhang, H. Li, B. Xu, X. Li, S. Liu, Y. Zhang and J. Xu, *J. Phys. Chem. B*, 2011, **115**, 7606–7611.
- 15 J. Luo, Z. Xie, J. W. Y. Lam, L. Cheng, H. Chen, C. Qiu, H. S. Kwok, X. Zhan, Y. Liu, D. Zhu and B. Z. Tang, *Chem. Commun.*, 2001, 1740–1741.
- 16 (a) G. Ruggeri and A. Pucci, *Chem. Soc. Rev.*, 2013, **42**, 857–870; (b) M. Montalti, L. Prodi, E. Rampazzo and N. Zaccheroni, *Chem. Soc. Rev.*, 2014, **43**, 4243–4268.
- 17 M. Chen, L. Li, H. Nie, J. Tong, L. Yan, B. Xu, J. Z. Sun, W. Tian, Z. Zhao, A. Qin and B. Z. Tang, *Chem. Sci.*, 2015, **6**, 1932–1937.
- 18 (a) Z. Zhao, S. Chen, X. Shen, F. Mahtab, Y. Yu, P. Lu, J. W. Y. Lam, H. S. Kwok and B. Z. Tang, *Chem. Commun.*,



- 2010, **46**, 686–688; (b) Z. Zhao, P. Lu, Z. Wang, C. Y. K. Chan, H. H. Y. Sung, I. D. Williams, Y. Ma and B. Z. Tang, *Chem. Sci.*, 2011, **2**, 672–675.
- 19 S. Sasaki, Y. Niko, K. Igawa and G. Konishi, *RSC Adv.*, 2014, **4**, 33474–33477.
- 20 E. Hasegawa, S. Takizawa, T. Seida, A. Yamaguchi, N. Yamaguchi, N. Chiba, T. Takahashi, H. Ikeda and K. Akiyama, *Tetrahedron*, 2006, **62**, 6581–6588.
- 21 J. Dey and I. M. Warner, *J. Phys. Chem. A*, 1997, **101**, 4872–4878.
- 22 (a) D. S. Surry and S. L. Buchwald, *Chem. Sci.*, 2011, **2**, 27–50; (b) J. P. Wolfe, S. Wagaw and S. L. Buchwald, *J. Am. Chem. Soc.*, 1996, **118**, 7215–7216.
- 23 (a) M. G. Organ, M. Abdel-Hadi, S. Avola, I. Dubivik, N. Hadei, E. A. B. Kantchev, C. J. O'Brien, M. Sayah and C. Valente, *Chem. – Eur. J.*, 2008, **14**, 2443–2452; (b) Y. Suzuki, N. Fukui, K. Murakami, H. Yorimitsu and A. Osuka, *Asian J. Org. Chem.*, 2013, **2**, 1066–1071.
- 24 Y. Chung, B. F. Duerr, T. A. Mckelvey, P. Nanjappan and A. W. Czarnik, *J. Org. Chem.*, 1989, **54**, 1018–1032.
- 25 T. W. Campbell, V. E. McCoy, J. C. Kauer and V. S. Foldi, *J. Org. Chem.*, 1961, **26**, 1422–1426.
- 26 The only report relevant to **1,4-BPA** describes 1,4-diaminoanthracene and its UV-vis spectrum; see: E. A. Venediktov, V. A. Perfil'ev and B. D. Berezin, *Russ. J. Phys. Chem.*, 1985, **59**, 1824–1826.
- 27 The synthesis of a dimethylamine analogue of **1,5-BPA** was reported, but unfortunately photophysical properties were not described; see: T. Harada, T. Hiramatsu and T. Yamaji, *Synth. Commun.*, 1981, **11**, 379–384.
- 28 The synthesis of **1,8-BPA** was previously reported, albeit in the absence of any photophysical measurements. UV-vis spectra are only available for 1,8-diaminoanthracene; see: (a) I. P. Beletskaya, A. G. Bessemertnykh, A. D. Averin, F. Denat and R. Guillard, *Eur. J. Org. Chem.*, 2005, 281–305; (b) A. Dahan, T. Ashkenazi, V. Kuznetsov, S. Makievski, E. Drug, L. Fadeev, M. Bramson, S. Schokory, E. R. Kemelmakher and M. Gozin, *J. Org. Chem.*, 2007, **72**, 2289–2296.
- 29 (a) The synthesis of a dimethylamine analogue of **2,6-BPA** was reported, albeit in the absence of any photophysical measurements. UV-vis spectra are only available for 2,6-diaminoanthracene; see: F. Keller and C. Ruchardt, *J. Prakt. Chem.*, 1998, **340**, 642–648; (b) W. S. Trahanovsky, J. L. Tunkel, J. C. Thoen and Y. Wang, *J. Org. Chem.*, 1995, **60**, 8407–8409.
- 30 The synthesis of a dimethylamine analogue of **9,10-BPA** was reported, and the photophysical properties were described as “non-fluorescent”; see: ref. 20.
- 31 (a) M. P. Aldred, C. Li and M. Q. Zhu, *Chem. – Eur. J.*, 2012, **18**, 16037–16045; (b) A. J. Qin, L. Tang, J. W. Y. Lam, C. K. W. Jim, Y. Yu, H. Zhao, J. Z. Sun and B. Z. Tang, *Adv. Funct. Mater.*, 2009, **19**, 1891–1900.
- 32 K. T. Wong, Y. Y. Chien, R. T. Chen, C. F. Wang, Y. T. Lin, H. H. Chiang, P. Y. Hsieh, C. C. Wu, C. H. Chou, Y. O. Su, G. H. Lee and S. M. Peng, *J. Am. Chem. Soc.*, 2002, **124**, 11576–11577.
- 33 (a) A. Wakamiya, K. Mori and S. Yamaguchi, *Angew. Chem., Int. Ed.*, 2007, **46**, 4273–4276; (b) M. Shimizu, Y. Takeda, M. Higashi and T. Hiyama, *Angew. Chem., Int. Ed.*, 2009, **355**, 3707–3710; (c) R. Yoshii, A. Hirose, K. Tanaka and Y. Chujo, *J. Am. Chem. Soc.*, 2014, **136**, 18131–18139.
- 34 R. Hu, E. Lager, A. A. Aguilar, J. Liu, J. W. Y. Lam, H. H. Y. Sung, I. D. Williams, Y. Zhong, K. S. Wong, E. P. Cabrera and B. Z. Tang, *J. Phys. Chem. C*, 2009, **113**, 15845–15853.
- 35 (a) H. Deng, B. Liu, C. Yang, G. Li, Y. Zhuang, B. Li and X. Zhu, *RSC Adv.*, 2014, **4**, 62021–62029; (b) C. Wu, B. Bull, C. Szymanski, K. Christensen and J. McNeil, *ACS Nano*, 2008, **2**, 2415–2423.
- 36 (a) L. Schermelleh, R. Heintzmann and H. Leonhardt, *J. Cell Biol.*, 2010, **190**, 165–175; (b) B. Huang, *Curr. Opin. Chem. Biol.*, 2010, **14**, 10–14; (c) E. A. J. Erijiman and T. M. Jovin, *Nat. Biotechnol.*, 2003, **21**, 1387–1395.
- 37 (a) H. Schill, S. Nizamov, F. Bottanelli, J. Bierwagen, V. N. Belov and S. W. Hell, *Chem. – Eur. J.*, 2013, **19**, 16556–16565; (b) Y. Yang, M. Lowry, X. Xu, J. O. Escobedo, M. S. Vazquez, L. Wong, C. M. Schowalter, T. M. Jensen, F. R. Fronczek, I. M. Warner and R. M. Strongin, *Proc. Natl. Acad. Sci. U. S. A.*, 2008, **105**, 8829–8834; (c) P. Horvath, P. Sebej, T. Solomek and P. Klan, *J. Org. Chem.*, 2015, **80**, 1299–1311; (d) J. F. Araneda, W. E. Piers, B. Heyne, M. Parvez and R. McDonald, *Angew. Chem., Int. Ed.*, 2011, **50**, 12214–12217; (e) Y. Y. Wu, Y. Chen, G. Z. Gou, W. H. Mu, X. J. Lv, M. L. Du and W. F. Fu, *Org. Lett.*, 2012, **14**, 5226–5229.
- 38 (a) D. W. Schlosser and M. R. Philpott, *Chem. Phys.*, 1980, **49**, 181; (b) H. Yamagata, J. Norton, E. Hontz, Y. Oliver, D. Beljonne, J. L. Bredas, R. J. Silbey and F. C. Spano, *J. Chem. Phys.*, 2011, **134**, 204703.
- 39 (a) R. Steiger, R. Pugin and J. Heier, *Colloids Surf., B*, 2009, **74**, 484–491; (b) H. B. Fu, Y. Q. Wang and J. N. Yao, *Chem. Phys. Lett.*, 2000, **322**, 327–332.
- 40 A. Sillen and Y. Engelborghs, *Photochem. Photobiol.*, 1998, **67**, 475–486.
- 41 N. J. Turro, V. Ramamurthy and J. C. Scaiano, *Modern Molecular Photochemistry of Organic Molecules*, University Science Books, Sausalito, 2010, pp. 388–413.
- 42 (a) O. G. Reid, R. D. Pensack, Y. Song, G. D. Scholes and G. Rumbles, *Chem. Mater.*, 2014, **26**, 561–571; (b) A. K. Thomas, J. A. Garcia, J. U. Sanchez, J. Gao and J. K. Grey, *ACS Nano*, 2014, **8**, 10559–10568.
- 43 (a) C. Deng, Q. Peng, Y. Niu and Z. Shuai, *J. Comput. Chem.*, 2012, **33**, 1862–1869; (b) Q. Wu, Q. Peng, Y. Niu, X. Gao and Z. Shuai, *J. Phys. Chem. A*, 2012, **116**, 3881–3888; (c) J. Mei, Y. Hong, J. W. Y. Lam, A. Qin, Y. Tang and B. Z. Tang, *Adv. Mater.*, 2014, **26**, 5429–5479.
- 44 A. P. Darmanyan, *Chem. Phys. Lett.*, 1982, **91**, 396–400.
- 45 X. Zhang, Z. Chi, B. Xu, C. Chen, X. Zhou, Y. Zhang and S. Liu, *J. Mater. Chem.*, 2012, **22**, 18505–18513.
- 46 M. Montalti, A. Credi, L. Prodi and M. T. Gandolfi, *Handbook of Photochemistry*, Taylor & Francis, Boca Raton, 2006, pp. 536–540.



- 47 A. Iida and S. Yamaguchi, *Chem. Commun.*, 2009, 3002–3004.
- 48 (a) E. L. Lippert, *Organic Molecular Photophysics*, Wiley, New York, 1975; (b) N. Mataga and T. Kubota, *Molecular Interaction and Electronic Spectra*, Dekker, New York, 1970; (c) A. Fin, A. V. Jentzsch, N. Sakai and S. Matile, *Angew. Chem., Int. Ed.*, 2012, **51**, 12736–12739.
- 49 (a) K. Rotkiewicz and Z. R. Grabowski, *Trans. Faraday Soc.*, 1969, **65**, 3263–3278; (b) M. Tichy and R. Zahradnik, *J. Phys. Chem.*, 1969, **73**, 534–544.
- 50 S. Watanabe, K. Kumagai, M. Hasegawa, M. Kobayashi, J. Okubo, T. Yoshinaga, H. Hiratsuka and T. Hoshi, *Bull. Chem. Soc. Jpn.*, 2000, **73**, 1783–1789.
- 51 R. P. Steiner and J. Michl, *J. Am. Chem. Soc.*, 1978, **100**, 6861–6867.
- 52 S. Zilberg, Y. Haas and S. Shaik, *J. Chem. Phys.*, 1995, **99**, 16558–16565.
- 53 N. Nijegorodov and R. Mabbs, *Spectrochim. Acta, Part A*, 2000, **56**, 2157–2166.
- 54 Z. R. Grabowski, K. Rotkiewicz and W. Rettig, *Chem. Rev.*, 2003, **103**, 3899–4031.
- 55 T. Tahara and H. Hamaguchi, *Chem. Phys. Lett.*, 1994, **217**, 369.
- 56 (a) A.-D. Gorse and M. Pesquer, *J. Phys. Chem.*, 1995, **99**, 4039–4049; (b) C. J. Jodicke and H. P. Luthi, *J. Am. Chem. Soc.*, 2003, **125**, 252–264; (c) T. Taniguchi, J. Wang, S. Irie and S. Yamaguchi, *Dalton Trans.*, 2013, **42**, 620–624.
- 57 (a) J. Li, Y. Qian, L. Xie, Y. Yi, W. Li and W. Huang, *J. Phys. Chem. C*, 2015, **119**, 2133–2141; (b) Y. Qian, M.-M. Cai, L. H. Xie, G.-Q. Yang, S.-K. Wu and W. Huang, *Chem-PhysChem*, 2011, **12**, 397–404.
- 58 L. R. Lakowicz, *Principles of Fluorescence Spectroscopy Third Edition*, Springer Science+Business Media, LLC, New York, 2006, p. 54.
- 59 G. Kortüm, W. Braun and G. Herzog, *Angew. Chem., Int. Ed.*, 1963, **2**, 333–341.
- 60 B. Chu, *Laser Light Scattering Basic Principles and Practice Second Edition*, Dover Publications, Mineola, 2007.
- 61 S. K. Lee, W. J. Yang, J. J. Choi, C. H. Kim, S.-J. Jeon and B. R. Cho, *Org. Lett.*, 2005, **7**, 323–326.
- 62 S. E. Wheeler, A. J. McNeil, P. Muller, T. M. Swager and K. N. Houk, *J. Am. Chem. Soc.*, 2010, **132**, 3304–3311.
- 63 A. Richardson, K. R. Brower and E. D. Amstutz, *J. Org. Chem.*, 1956, **21**, 890–891.

