



Cite this: *React. Chem. Eng.*, 2024, 9, 888

## Interaction of SO<sub>2</sub> with a Cu–Mn oxide oxygen carrier during chemical looping with oxygen uncoupling†

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Chemical looping with oxygen uncoupling (CLOU) is a variant of the chemical looping combustion (CLC) process, where fuel combustion occurs *via* metal oxides, known as oxygen carriers. The purpose of this study is to investigate how SO<sub>2</sub> interacts with a Cu–Mn oxide oxygen carrier during CH<sub>4</sub> combustion in the CLOU process. The oxygen carrier was reduced *via* CH<sub>4</sub>–N<sub>2</sub> gas mixture with and without SO<sub>2</sub> in a batch fluidized-bed reactor. The results indicate that SO<sub>2</sub> does not affect the oxygen release capacity of a Cu–Mn oxide. However, it negatively affects the CH<sub>4</sub> conversion by increasing CO formation as well as the unreacted CH<sub>4</sub> amount. This adverse effect becomes more pronounced as the SO<sub>2</sub> concentration increases. Increasing the temperature from 850 to 950 °C improves the conversion of CH<sub>4</sub>. During a reduction cycle, SO<sub>2</sub> oxidizes to SO<sub>3</sub>, creating a competing demand for oxygen between fuel and SO<sub>2</sub>, and lowers CH<sub>4</sub> conversion if there is not enough oxygen for both reactions. Hence, decreasing the feed CH<sub>4</sub> concentration with a fixed oxygen carrier amount improves CH<sub>4</sub> conversion in the presence of SO<sub>2</sub>. Additionally, characterization of SO<sub>2</sub>-exposed oxygen carriers shows that sulfate species form on the surface, possibly causing a further reduction in CH<sub>4</sub> conversion. However, the presence of SO<sub>2</sub> does not cause a permanent deactivation of the particles and they can be fully regenerated during the oxidation cycle.

Received 22nd September 2023,  
 Accepted 21st December 2023

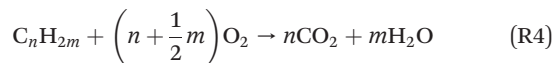
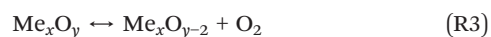
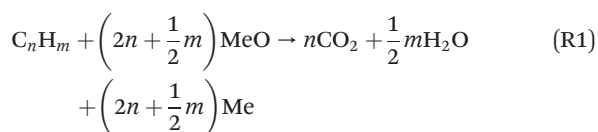
DOI: 10.1039/d3re00498h

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## Introduction

Fossil-fuel combustion accounts for about 60% of total electricity in the U.S. despite increased renewable energy resources.<sup>1</sup> Burning fossil fuels generates a considerable amount of CO<sub>2</sub>, a significant greenhouse gas. Carbon capture and storage (CCS) is a promising technology to limit CO<sub>2</sub> concentration in the atmosphere. Since a high purity CO<sub>2</sub> stream is essential to ensure the economic viability of CCS technologies, several carbon-capture methods have been proposed and investigated in recent years.<sup>2–5</sup> Of these methods, chemical looping combustion (CLC) is a promising technology because of its inherent ability to produce N<sub>2</sub>-free CO<sub>2</sub> during fuel combustion while bypassing an expensive gas separation method. In CLC, fuel oxidation occurs *via* a metal oxide, also referred to as an oxygen carrier, instead of air, *via* reaction (R1), therefore avoiding direct contact between fuel and air. As a result, after steam condensation, a highly concentrated CO<sub>2</sub> stream is obtained. The reduced oxygen

carrier is regenerated by air *via* (R2) before being transferred back to the fuel reactor.<sup>6–12</sup>



The oxidation by which a metal oxide's lattice oxygen reacts with the fuel is the primary distinguishable characteristic of chemical looping technologies. This phenomenon may occur in two ways: i) surface reactions between the fuel and lattice oxygen, (R1), and ii) reaction with gaseous oxygen, (R4), from metal oxide decomposition, (R3), at suitable temperature and oxygen partial pressures. The latter mechanism, where fuel reacts with gaseous oxygen like the traditional combustion reaction, is commonly referred to as chemical looping with oxygen uncoupling (CLOU). The CLOU mechanism is

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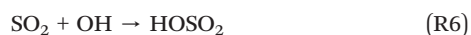
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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3re00498h>

particularly beneficial for solid fuel utilization due to a faster gas–solid reaction rate relative to a solid–solid reaction in CLC.<sup>7,13–22</sup>

Early CLOU research identified monometallic CuO/Cu<sub>2</sub>O, Mn<sub>2</sub>O<sub>3</sub>/Mn<sub>3</sub>O<sub>4</sub> oxide-systems as suitable CLOU oxygen carriers. However, these oxides have some limitations like sintering of copper oxides and slow oxidation kinetics of Mn<sub>3</sub>O<sub>4</sub> to Mn<sub>2</sub>O<sub>3</sub>.<sup>23</sup> Therefore, to overcome this limitation, copper and manganese oxides are combined with inert oxides such as TiO<sub>2</sub>, ZrO<sub>2</sub>, SiO<sub>2</sub>, or active metal oxides such as NiO, CaO, and Fe<sub>2</sub>O<sub>3</sub>.<sup>14,24–27</sup> One such oxygen carrier is bimetallic Cu–Mn oxide, which utilizes the oxygen uncoupling characteristics of CuO and Mn<sub>2</sub>O<sub>3</sub> and negates the disadvantages of constituent oxides.<sup>28</sup> It transitions between spinel Cu<sub>x</sub>Mn<sub>3–x</sub>O<sub>4</sub> and CuMnO<sub>2</sub> during oxygen uncoupling that provides a considerable amount of gaseous O<sub>2</sub> that results in high reactivity with the fuel. Furthermore, CuMnO<sub>2</sub> can also react with the fuel *via* heterogeneous reactions (CLC) to form Cu and MnO.<sup>29–31</sup> Previous studies employing a Cu–Mn oxide achieved high combustion efficiency for both solid and gaseous fuels.<sup>28,32–37</sup> However, the investigations mainly involved sulfur-free or low-sulfur fuels to highlight the redox reactivity of the oxygen carrier. Although one study tested high-sulfur coal (5.2 wt% S) with a Cu–Mn oxygen carrier, it did not investigate the potential impact that sulfur contamination may have on the carrier.<sup>38</sup> Given the variable presence of sulfur in fuels such as coal (200–10 000 ppm H<sub>2</sub>S in coal-derived synthesis gas) and natural gas (up to 20 ppm H<sub>2</sub>S), the oxygen carrier will most likely be exposed to sulfur-containing species in the fuel reactor.<sup>39,40</sup>

Sulfur is found in coal in various forms, including sulfides, sulfates, and organic sulfur compounds and most part of the coal-S is released during combustion from sulfides and organic bound sulfur.<sup>41</sup> At temperatures higher than 1000 °C and oxygen-rich conditions, SO<sub>2</sub> is the thermodynamically favored species. H<sub>2</sub>S can be formed in the gas phase or released from coal-S under substoichiometric conditions. At lower temperatures, the equilibrium shifts towards SO<sub>3</sub> and typically 0.1 to 1% of the SO<sub>2</sub> is oxidized to SO<sub>3</sub> during air combustion *via* reactions (R5)–(R7) in the gas phase.<sup>42</sup>



SO<sub>3</sub> formation is influenced by the temperature profile, residence time, concentrations of SO<sub>2</sub> and O<sub>2</sub>, fly ash composition and presence of catalysts *e.g.* Fe<sub>2</sub>O<sub>3</sub>.<sup>43</sup> In the CLOU process, with the presence of the oxygen carrier acting as an oxidizing catalyst, SO<sub>2</sub> could be oxidized to SO<sub>3</sub> heterogeneously. Due to the likelihood of the interactions of sulfur species and the oxygen carrier, understanding how sulfur species affect the oxygen carrier's performance is

crucial to developing the chemical looping technology. Nevertheless, studies looking into the impacts of sulfur on Cu and Mn-based oxygen carriers during the CLOU process are limited. A study investigating high-sulfur lignite combustion in a continuous CLOU reactor with CuO–MgAl<sub>2</sub>O<sub>4</sub> observed SO<sub>2</sub> as the major sulfur species in the effluent. In this study, the sulfur mass balance closure was ~80 wt%, with the rest remaining unaccounted. The authors assumed that the remaining sulfur was accumulated on the oxygen carrier; however, they did not observe a decrease in the reactivity of the oxygen carrier due to the presence of sulfur.<sup>44</sup> In a similar CLOU study with a Cu-based oxygen carrier (CuO–Fe<sub>2</sub>O<sub>3</sub>–MgAl<sub>2</sub>O<sub>4</sub>), SO<sub>2</sub> was also observed as the final sulfur product.<sup>45</sup> They observed no decrease in the reactivity of the oxygen carrier when exposed to sulfurous fuel; however, the oxygen uncoupling capacity decreased ~10% after 35 hours of operation. Around 15 wt% of the total sulfur was unaccounted for in the sulfur mass balance. Through a controlled TGA study followed by SEM-EDX analysis, (CuO)·(CuSO<sub>4</sub>) formation was observed at 930 °C even though it is thermodynamically unstable above 700 °C. The authors speculated that combination of different oxides in the oxygen carrier causes a synergic effect making its reaction affinity with SO<sub>2</sub> different from those of the individual oxides.<sup>45</sup> Besides Cu-based oxides, a study with a Mn–Si oxide investigated sulfur effects by introducing 5000 ppm SO<sub>2</sub> with CH<sub>4</sub>.<sup>46</sup> The oxygen carrier displayed no appreciable decrease in reactivity or oxygen uncoupling capacity at a maximum exposure of 5000 ppm SO<sub>2</sub>.

On the other hand, Cu and Mn-based oxides tend to form sulfides in the CLC process depending on the operating conditions.<sup>39,47–51</sup> For Cu-based oxygen carriers, thermodynamic calculations show that the possibility of sulfide formation is high at oxygen-deficient conditions, while it is less likely to form at oxygen-rich conditions. For a Mn-based oxide, MnSO<sub>4</sub> formation is thermodynamically possible at both oxygen-deficient and rich conditions.<sup>49</sup> A CLOU fuel reactor usually operates in an oxygen-rich environment which decreases the possibility of sulfide formation.<sup>52</sup> However, local regions with high reducing potentials may promote sulfide formation.

Given the prospects of Cu–Mn oxide as an oxygen carrier in CLOU and the potential probability of sulfur interaction with constituent elements, it is necessary to examine the influence of sulfur on this oxygen carrier. Therefore, the objective of this study is to evaluate how sulfur affects the Cu–Mn oxygen carrier in a CLOU process. The study simulates coal combustion in a CLOU process by introducing CH<sub>4</sub> as fuel since it is released during the volatilization stage of coal combustion. Since SO<sub>2</sub> is the major sulfur species formed during solid fuel combustion in CLOU,<sup>45</sup> it was added as a sulfur impurity. Effects of SO<sub>2</sub> on the oxygen carrier's reactivity with CH<sub>4</sub> was investigated in a laboratory-scale fluidized bed reactor. Finally, Cu–Mn oxygen carrier particles were characterized before and after SO<sub>2</sub> exposure to determine its interaction with SO<sub>2</sub>.

## Experimental

### Materials

The bi-metallic Cu–Mn oxide, synthesized by the incipient wetness impregnation method,<sup>53</sup> consists of 34 wt% CuO and 66 wt% Mn<sub>3</sub>O<sub>4</sub>. Before impregnation, Mn<sub>3</sub>O<sub>4</sub> was mixed with graphite (5 wt% of Mn<sub>3</sub>O<sub>4</sub>) and sintered at 950 °C for one hour. A certain amount of copper nitrate solution (60–80 °C) corresponding to the pore volume of Mn<sub>3</sub>O<sub>4</sub> was added to the sintered particles. The impregnated samples were dried at 70 °C overnight and were calcined at 600 °C in an air atmosphere for two hours. The desired loading was achieved by successive 3–4 impregnation steps followed by drying and calcination. Finally, the particles were heated at 950 °C for 6 hours at 10 °C min<sup>−1</sup> and sieved to a size of 150–300 μm. The final oxygen carrier obtained is referred to as Cu34Mn66 in this study.

### Characterization of the oxygen carrier

The powder X-ray diffraction (XRD) method (Rigaku Ultima-III; 30 kV and 40 mA) with Cu-Kα radiation was used to determine the crystalline structure of the oxygen carrier. The pattern was collected in the range of 10 to 90°. The X-ray photoelectron spectroscopy (XPS) analysis was conducted with an AXIS Supra by Kratos Analytical (300 W, 20 mA) equipped with a hemispherical electron analyzer and monochromatic Al-Kα source to analyze the surface chemical composition. Survey and high-resolution spectra of elements were acquired with 160 and 40 eV pass energies, respectively. The binding energies were calculated referencing the C 1s peak energy of 284.8 eV. Temperature programmed reduction (TPR) *via* hydrogen was performed using a Micromeritics 3Flex Surface and Catalyst Analyzer. A 10% H<sub>2</sub>/Ar mixture of 30 ml min<sup>−1</sup> flow rate was used to reduce 60 mg particles. Before reduction, particles were pretreated at 200 °C in N<sub>2</sub> for 1 hour, then cooled to 50 °C. The temperature was ramped up to 800 °C at a heating rate of 10 °C min<sup>−1</sup>. Isopropanol and liquid N<sub>2</sub> slurry were used to condense the water formed during the process. A thermal conductivity detector (TCD) was used to measure the H<sub>2</sub> consumption signal. Thermogravimetric analysis (TGA) was utilized to determine the oxygen release capacity of Cu34Mn66 particles. Netzsch TG 209 F1 Libra was used to analyze the oxygen release capacity at the following isothermal conditions: 850 °C, 900 °C, and 950 °C. Around 7–9 mg sample was reduced in N<sub>2</sub> and oxidized in air at a 180 ml min<sup>−1</sup> flow rate. The oxygen uncoupling capacity ( $R_o$ ) was determined using eqn (E1):

$$R_o = (m_{ox} - m_{red})/m_{ox} \quad (E1)$$

Here  $m_{ox}$  is the weight of the fully oxidized sample and  $m_{red}$  is the weight of the reduced sample in TGA.

### Experimental set-up

A quartz fluidized bed reactor was used to investigate the reactivity of the oxygen carrier with CH<sub>4</sub> in the presence of

SO<sub>2</sub> at CLOU conditions. The experimental set-up simulates the gas–solid contact conditions relevant for a large-scale chemical looping combustion process. Instead of two separate reactors, fuel and air feeds into a single reactor are alternated to create the reduction and oxidation environments in a batch manner. Fig. 1 presents the schematic of the experimental set-up used in this work.

The reactor has a 2 cm inner diameter at the lower section and a 2.5 cm inner diameter section at the top. The increase in diameter at the top section lowers the gas velocity to prevent the smaller particles to leave the reactor. A porous quartz disc inside the reactor supports the particles and allows the gas to flow through the particle bed. A Thermcraft furnace surrounds the reactor to heat it up to the desired temperature. An Omega PX-409 differential pressure transducer measures the pressure fluctuations across the reactor bed and the quartz disc, indicating the fluidization behavior of the particles. For all the experiments, the Cu34Mn66 oxygen carrier inventory was 4 g to ensure there is enough oxygen released for fuel conversion. Experiments were performed at various conditions by changing the temperature and concentrations of CH<sub>4</sub> and SO<sub>2</sub> as shown in the test matrix in Table 1. All the experiments were repeated three times to ensure the reproducibility. Each cycle consisted of a reduction–purge–oxidation pattern. Reduction of the oxygen carrier occurred *via* a reducing mixture of either CH<sub>4</sub>–N<sub>2</sub> or CH<sub>4</sub>–SO<sub>2</sub>–N<sub>2</sub> at a flowrate of 600 ml min<sup>−1</sup>. Duration of the reduction period was 1 minute for most of the cases. After the reduction cycle, inert N<sub>2</sub> was flowed through the reactor for around 1 minute to purge the reactor, followed by 700 ml min<sup>−1</sup> air flow to reoxidize the particles. The selected reduction and oxidation flow rates, respectively, correspond to about 2–8 and 2–9 times the theoretical minimum fluidization velocity ( $U_{mf}$ ). The  $U_{mf}$  value is calculated according to the correlations of Kunii and Levenspiel.<sup>54</sup> Flue gas species (CO, CO<sub>2</sub>, CH<sub>4</sub>, and O<sub>2</sub>) were

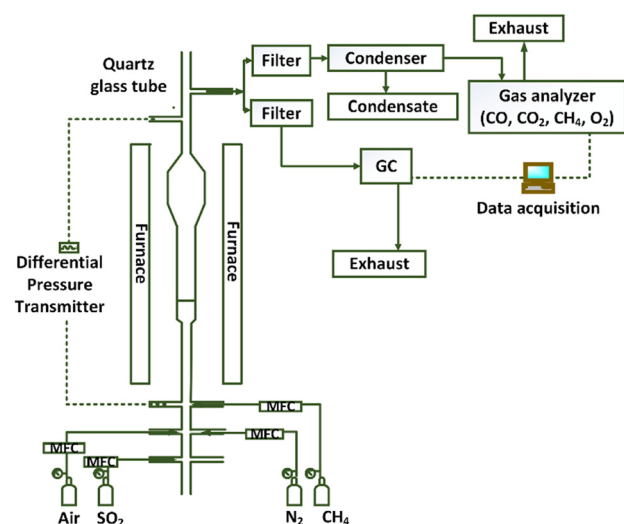


Fig. 1 Schematic of the experimental set-up.

**Table 1** Experimental conditions

Case	Oxygen carrier amount (g)	Temperature (°C)	SO <sub>2</sub> concentration (ppm)	CH <sub>4</sub> concentration (%)	Reducing time (min)
1	4	900	0, 1000, 3000, 5000	9	1
2	4	900	0, 5000	3, 6, 9	1
3	4	900	5000	9	1, 3, 8
4	4	850, 900, 950	0, 5000	9	1

analyzed at a dry basis by a continuous Rosemount X-stream enhanced analyzer to obtain the reduction profiles for each cycle. H<sub>2</sub> and S species were measured on a wet basis by a gas chromatograph (SRI MG#5 GC); therefore, eliminating the possibility of losing sulfur in condensed water. With a GC, the analysis of each sample takes 8 minutes, which makes it impossible to obtain a continuous profile of SO<sub>2</sub> in a single experiment. Therefore, fresh particles were exposed to different reduction periods in separate experiments to obtain the SO<sub>2</sub> concentration profile against time for a certain cycle. The GC data was collected at the end of each reducing period. So, each data point shown in the concentration profiles presented in the results section represents data from a separate experiment. For the experiments conducted to obtain the SO<sub>2</sub> concentration profiles, the reduction cycle was extended to 8 minutes, instead of 1 min, to see if prolonged exposure to SO<sub>2</sub> results in sulfur accumulation on the oxygen carrier. Additionally, the salt method was employed to qualitatively detect SO<sub>3</sub> in the gas phase as opposed to quantitatively measuring its concentration. This analysis was performed by flowing the fluidized bed reactor effluent from a specific reduction cycle through a small packed-bed reactor containing 1 g of sodium chloride (NaCl) without condensing the water.<sup>55–57</sup> The temperature of the sample line from the reactor to the packed bed was maintained at 180 °C, while the reactor temperature was maintained at 200 °C. In the presence of H<sub>2</sub>O vapor at this temperature, SO<sub>3</sub> reacts with the salt and form sodium sulfate.<sup>57</sup> The exposed salt sample was titrated by barium perchlorate (Ba(ClO<sub>4</sub>)<sub>2</sub>) solution to detect the presence of SO<sub>3</sub>.<sup>56</sup>

### Data evaluation

Reduction performance of the Cu34Mn66 was evaluated using total conversion of CH<sub>4</sub> and yields of CO<sub>2</sub> and CO. Total CH<sub>4</sub> conversion considers formation of both CO and CO<sub>2</sub> as calculated *via* eqn (E2).

$$\text{CH}_4 \text{ conversion (\%)} = \frac{[(X_{\text{CO}} + X_{\text{CO}_2}) / (X_{\text{CO}} + X_{\text{CO}_2} + X_{\text{CH}_4})] \times 100}{\quad} \quad (\text{E2})$$

CO<sub>2</sub> and CO yields were calculated *via* eqn (E3) and (E4).

$$\text{CO}_2 \text{ yield (\%)} = \frac{[X_{\text{CO}_2}] / (X_{\text{CO}} + X_{\text{CO}_2}) \times 100}{\quad} \quad (\text{E3})$$

$$\text{CO yield (\%)} = \frac{[X_{\text{CO}}] / (X_{\text{CO}} + X_{\text{CO}_2}) \times 100}{\quad} \quad (\text{E4})$$

Here,  $X_i$  is the molar fraction of the carbon gas species in the reactor outlet stream, with  $i$  denoted as CO, CO<sub>2</sub>, or unreacted CH<sub>4</sub>. The carbon mass balance closure for the 1 minute reduction period is  $0.98 \pm 0.013$  SE. From the sulfur balance around the reactor, the SO<sub>2</sub> loss is calculated by eqn (E5).

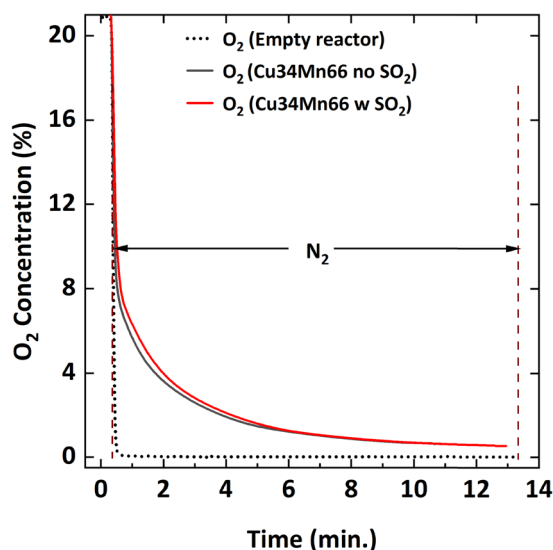
$$\text{SO}_2 \text{ loss (mol\%)} = [1 - (X_{\text{SO}_2, \text{out}} \cdot F_{\text{out}}) / (X_{\text{SO}_2, \text{in}} \cdot F_{\text{in}})] \times 100 \quad (\text{E5})$$

Here,  $X_{\text{SO}_2, \text{out}}$  is the molar fraction of SO<sub>2</sub> in the flue gas as measured by the GC analyzer and  $X_{\text{SO}_2, \text{in}}$  is the inlet molar fraction of SO<sub>2</sub>.  $F_{\text{out}}$  is the calculated outlet molar flow rate,  $F_{\text{in}}$  is the inlet molar flow rate of the reducing gas. The loss of SO<sub>2</sub> can occur from sulfur accumulation on the Cu34Mn66 particles, SO<sub>3</sub> formation, or both.

## Results

### Effect of SO<sub>2</sub> on oxygen uncoupling behavior of Cu34Mn66

Fig. 2 depicts the oxygen concentration profile during oxygen uncoupling in N<sub>2</sub> at 900 °C with and without 5000 ppm of SO<sub>2</sub>. Oxygen concentration measurement in an empty reactor is also reported to show that the N<sub>2</sub> flow clears the existing air in the reactor and the oxygen detected in presence of the oxygen carrier particles is the oxygen released from the particles. Without SO<sub>2</sub>, the oxygen concentration decreases over time, from 6% to 0.6% within 11 minutes. Similar



**Fig. 2** O<sub>2</sub> concentration profile during oxygen uncoupling of Cu34Mn66 particles in N<sub>2</sub> at 900 °C with and without 5000 ppm SO<sub>2</sub>.

behavior was observed in previous literature.<sup>36</sup> In the presence of SO<sub>2</sub>, the oxygen concentration is slightly higher until about 2.5%; then, the profile is the same. However, the presence of SO<sub>2</sub> does not affect the total oxygen uncoupling capacity of the particles after 20 cycles of SO<sub>2</sub> exposure. TGA results show that R0 remains at 4 wt%  $\pm$  0.02 wt% for the three cases tested: fresh oxidized particles, particles reduced for 20 cycles with 5000 ppm SO<sub>2</sub>, and particles reduced for 20 cycles without SO<sub>2</sub>.

### Chemical looping combustion of CH<sub>4</sub> without SO<sub>2</sub>

Initially, the Cu<sub>34</sub>Mn<sub>66</sub> particles were reduced with a mixture of 9% CH<sub>4</sub>-balance N<sub>2</sub> to determine its behavior in the absence of SO<sub>2</sub>. Fig. 3 shows the dry basis concentration profiles of the flue gas species with and without SO<sub>2</sub> addition during the 5th cycle at 900 °C. Corresponding CH<sub>4</sub> conversion and yields of CO<sub>2</sub> and CO are plotted in Fig. 4 and 5, respectively. In Fig. 3, during the pre-reduction N<sub>2</sub> purge cycle, O<sub>2</sub> concentration drops, but never reaches zero because the Cu<sub>34</sub>Mn<sub>66</sub> particles start releasing O<sub>2</sub> *via* the CLOU mechanism. Upon CH<sub>4</sub> addition without SO<sub>2</sub>, fuel oxidation occurs, producing CO<sub>2</sub> as the primary product and a small amount of CO. Oxidation can occur from two reactions: 1) CH<sub>4</sub> oxidation with the released O<sub>2</sub> in the gas phase and 2) direct reaction with the Cu<sub>34</sub>Mn<sub>66</sub> particles. No H<sub>2</sub> forms during reduction, as confirmed by the GC analysis. As the reduction proceeds, oxygen content of the Cu<sub>34</sub>Mn<sub>66</sub> particles depletes. Consequently, CO<sub>2</sub> concentration decreases while CO and unconverted CH<sub>4</sub> concentrations start to increase. Total CH<sub>4</sub> conversion is between 90–99% throughout the 1 minute reduction period as seen in Fig. 4. Slight amount of CO formation during the first 10 seconds results in a decrease in CO<sub>2</sub> as observed from the CO<sub>2</sub> yield

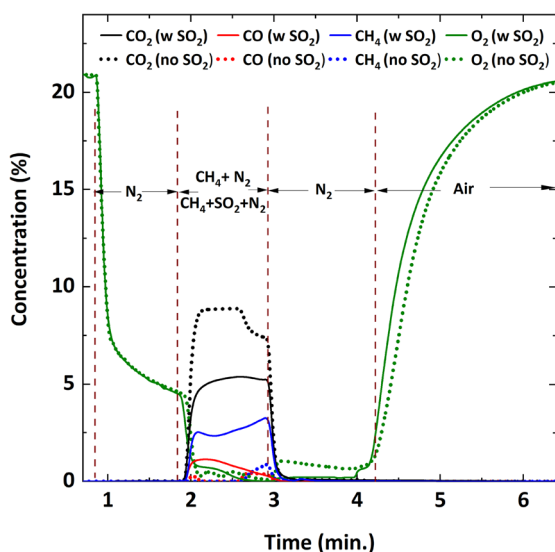


Fig. 3 Flue gas concentration profiles during the 5th cycle at 900 °C with and without 5000 ppm SO<sub>2</sub>. The reducing gas is 9% CH<sub>4</sub>-balance N<sub>2</sub> and 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub>. The oxidizing gas is air.

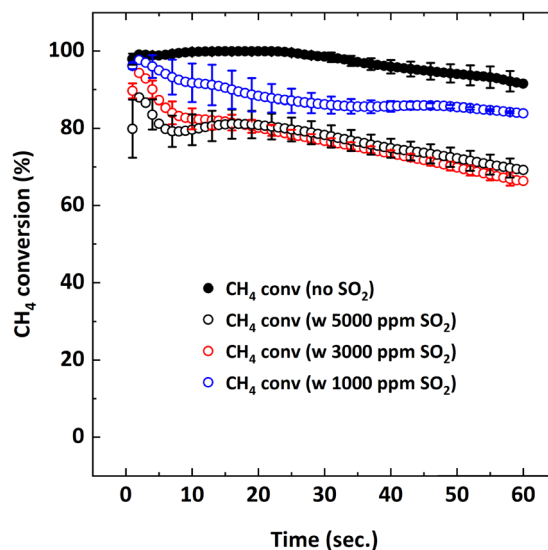


Fig. 4 CH<sub>4</sub> conversion as a function of reduction period during the 5th cycle at 900 °C with 9% CH<sub>4</sub>-X ppm SO<sub>2</sub>-balance N<sub>2</sub> where X: 0, 1000, 3000, 5000.

shown in Fig. 5. It is also worth noting that O<sub>2</sub> release still occurs from the Cu<sub>34</sub>Mn<sub>66</sub> particles (about 1%) during the subsequent inert period after stopping the reducing gas flow of 9% CH<sub>4</sub>-balance N<sub>2</sub> as seen in Fig. 3.

### Chemical looping combustion of CH<sub>4</sub> with SO<sub>2</sub>

Before starting the reduction cycle in the presence of SO<sub>2</sub>, a mixture of 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> was introduced to an empty reactor at 900 °C to evaluate the possibility of homogenous gas-phase reactions between CH<sub>4</sub> and SO<sub>2</sub>. No such gas-phase reactions were observed, as confirmed by the

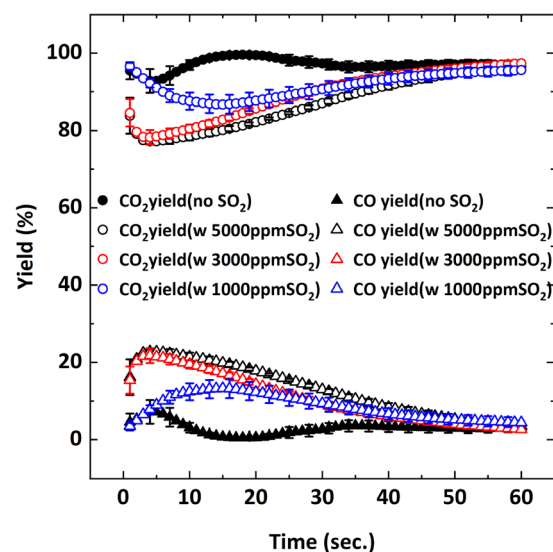


Fig. 5 CO<sub>2</sub> and CO yields as a function of reduction period during the 5th cycle at 900 °C with 9% CH<sub>4</sub>-X ppm SO<sub>2</sub>-balance N<sub>2</sub> where X: 0, 1000, 3000, 5000.

steady  $\text{CH}_4$  concentration profile. Then the reduction cycle was performed in the presence of the oxygen carrier particles with a mixture of 9%  $\text{CH}_4$ -X ppm  $\text{SO}_2$ -balance  $\text{N}_2$  (X: 1000, 3000, 5000 ppm). Fig. 3 depicts the measured flue gas concentration profiles with 5000 ppm  $\text{SO}_2$  addition during the 5th cycle at 900 °C. When 5000 ppm  $\text{SO}_2$  is added, CO and unreacted  $\text{CH}_4$  amounts are higher initially, compared to the  $\text{SO}_2$ -free reduction, resulting in a lower production of  $\text{CO}_2$ . Increasing the  $\text{SO}_2$  concentration from 1000 ppm to 5000 ppm lowers the conversion of  $\text{CH}_4$ , while increasing the CO yield resulting in a decrease in  $\text{CO}_2$  yield. Concentration profiles in Fig. 3 shows that negligible  $\text{O}_2$  release occurs from the particles during the subsequent  $\text{N}_2$  purge after reduction, unlike the case without  $\text{SO}_2$ , meaning the oxygen content is depleted when  $\text{SO}_2$  is present. Based on this observation, the presence of  $\text{SO}_2$  may either contribute to particle deactivation or the consumption of available oxygen during the reduction cycle.

An extended reduction cycle was also carried out to observe the impact of  $\text{SO}_2$  on  $\text{CH}_4$  combustion as a function of the reduction duration up to 9 minutes. Fig. 6 shows the profiles of the flue gas species obtained during complete reduction of Cu34Mn66 with and without 5000 ppm  $\text{SO}_2$ . In general, as the duration of the reduction cycle increases, the degree of reduction of the oxygen carrier particles also increases. Without  $\text{SO}_2$ ,  $\text{CO}_2$  concentration gradually decreases while the unreacted  $\text{CH}_4$  concentration increases. However, when  $\text{SO}_2$  is present,  $\text{CO}_2$  and  $\text{CH}_4$  concentrations follow a relatively steady profile between ~2.5–4.5 minutes. In addition, the complete reduction of Cu34Mn66 with  $\text{SO}_2$  takes slightly longer to complete than the 8 minutes required for  $\text{SO}_2$ -free reduction.

Analysis of post-combustion species with GC shows no formation of reduced sulfur gas species such as COS and  $\text{H}_2\text{S}$

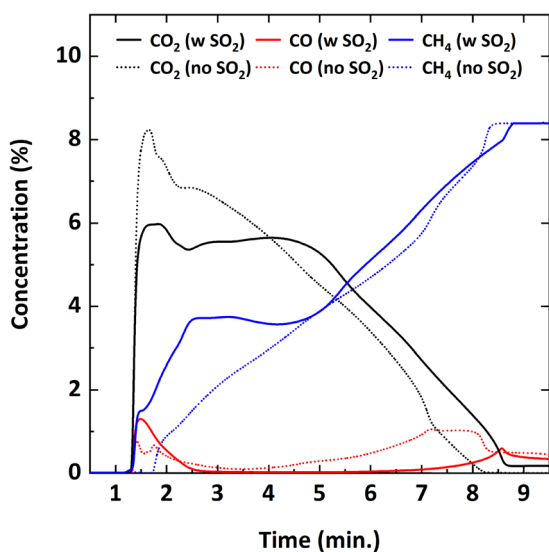


Fig. 6 Flue gas concentration profiles during complete reduction of Cu34Mn66 particles at 900 °C with 9%  $\text{CH}_4$ -balance  $\text{N}_2$  and 9%  $\text{CH}_4$ -5000 ppm  $\text{SO}_2$ -balance  $\text{N}_2$ .

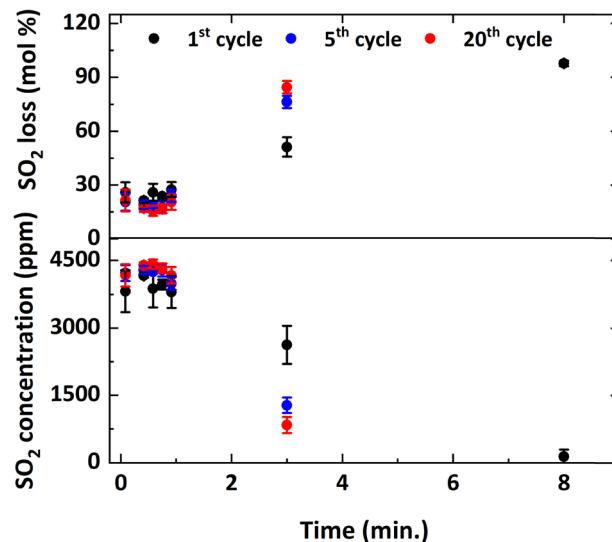


Fig. 7 Measured  $\text{SO}_2$  concentration profile and calculated  $\text{SO}_2$  loss (mol%) as a function of reduction period for the 1st, 5th, and 20th cycles at 900 °C with 9%  $\text{CH}_4$ -5000 ppm  $\text{SO}_2$ -balance  $\text{N}_2$ .

during oxygen carrier reduction *via* 9%  $\text{CH}_4$ -5000 ppm  $\text{SO}_2$ -balance  $\text{N}_2$ . Fig. 7 shows measured  $\text{SO}_2$  concentration profile along with calculated  $\text{SO}_2$  loss as a function of reducing time (up to 8 minutes) for the 1st, 5th, and 20th reduction cycles. For the 8-minute reduction, a single cycle was run due to observed particle agglomeration since the oxygen carrier is completely reduced. For the 1-minute reduction, about 25% of the  $\text{SO}_2$  feed is lost throughout and the increasing number of cycles does not affect the  $\text{SO}_2$  outlet concentration or the corresponding  $\text{SO}_2$  loss during this time. However, for the 3-minute reduction, increasing the cycle number increases the  $\text{SO}_2$  loss. When the reduction extent of the particles

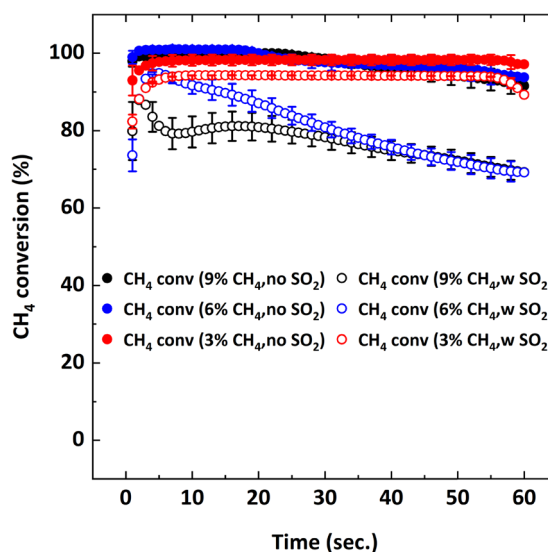


Fig. 8  $\text{CH}_4$  conversion as a function of reduction period during the 5th cycle at 900 °C with Y%  $\text{CH}_4$ -balance  $\text{N}_2$  and Y%  $\text{CH}_4$ -5000 ppm  $\text{SO}_2$ -balance  $\text{N}_2$  where Y: 3, 6, 9.

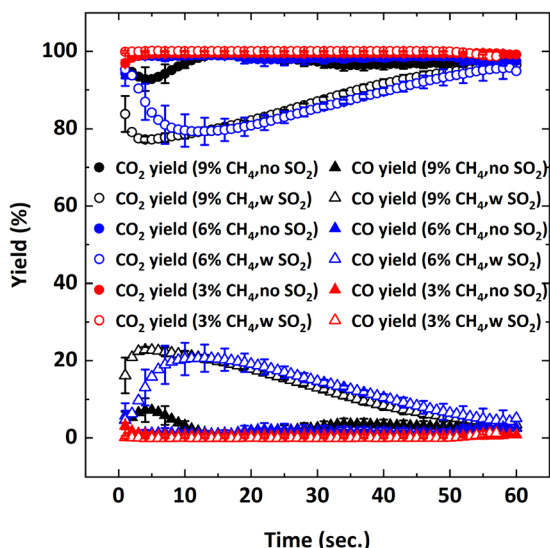


Fig. 9 CO<sub>2</sub> and CO yields as a function of reduction period during the 5th cycle at 900 °C with Y% CH<sub>4</sub>-balance N<sub>2</sub> and Y% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> where Y: 3, 6, 9.

increases, SO<sub>2</sub> loss also increases with almost no SO<sub>2</sub> left in the flue gas after 8 minutes of reduction.

To determine the likely cause of SO<sub>2</sub> loss in the gas phase, the salt method was employed to detect SO<sub>3</sub> if there is any in the gas phase. SO<sub>2</sub> could possibly be oxidized to SO<sub>3</sub> during the reduction cycle *via* either homogenous gas-phase or surface catalyzed reaction, resulting in the observed loss of SO<sub>2</sub> in the gas-phase. For the 1-minute and 3-minute reduction periods, SO<sub>3</sub> was observed during the 5th and 20th cycles, but not during the 1st cycle. However, for the 8-minute reduction it was observed during the 1st cycle.

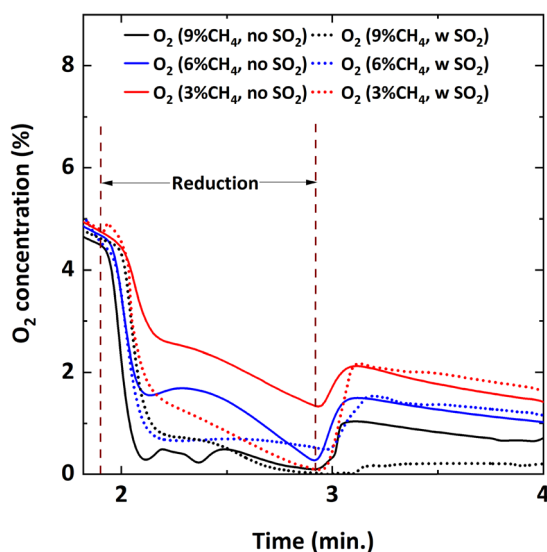


Fig. 10 O<sub>2</sub> concentration profile during the 5th cycle at 900 °C with Y% CH<sub>4</sub>-balance N<sub>2</sub> and Y% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> where Y: 3, 6, 9.

Since SO<sub>2</sub> is oxidized to SO<sub>3</sub>, in order to see if the availability of O<sub>2</sub> changes the effect of SO<sub>2</sub> on the oxygen carrier's reactivity, experiments were run with different CH<sub>4</sub> concentrations in the feed. Oxygen carrier particles were reduced with gas mixtures of Y% CH<sub>4</sub>-balance N<sub>2</sub> and Y% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> (Y: 3, 6, 9), and the corresponding CH<sub>4</sub> conversions and the yields of CO<sub>2</sub> and CO are provided in Fig. 8 and 9, respectively.

Without SO<sub>2</sub>, on average, 95–100% CH<sub>4</sub> conversion occurs and there is not much difference in conversion for all the CH<sub>4</sub> concentrations tested. As discussed before, CH<sub>4</sub> conversion decreases in the presence of SO<sub>2</sub>. The adverse outcome of SO<sub>2</sub> is inversely proportional to CH<sub>4</sub> concentration. For example, for 3% CH<sub>4</sub> in the feed, the effect of SO<sub>2</sub> on CH<sub>4</sub> conversion is minimal while the effect is more pronounced for the 9% CH<sub>4</sub> case. To explain this behavior, the oxygen concentration profiles were examined. Fig. 10 shows the O<sub>2</sub> concentration profiles for different CH<sub>4</sub> concentrations with and without 5000 ppm SO<sub>2</sub>.

Lowering the CH<sub>4</sub> feed concentration from 9% increases gaseous O<sub>2</sub> availability during the reduction cycle. In other words, for the lower CH<sub>4</sub> concentrations (6% and 3%), more O<sub>2</sub> is released from the oxygen carrier than the amount needed for the combustion reaction with CH<sub>4</sub>. However, when SO<sub>2</sub> is added, the availability of O<sub>2</sub> decreases and the O<sub>2</sub> concentration profiles become similar for all the CH<sub>4</sub> concentrations. This could be because the excess oxygen is being used up by SO<sub>2</sub> to become oxidized to SO<sub>3</sub>; therefore, indicating a competing oxygen demand between SO<sub>2</sub> and CH<sub>4</sub>. However, since Fig. 7 shows that SO<sub>2</sub> loss% increases with the reduction period, it can also be assumed that sulfur interacts with the reduced oxygen carrier. Post combustion analysis of the oxygen carrier is discussed later to investigate this hypothesis.

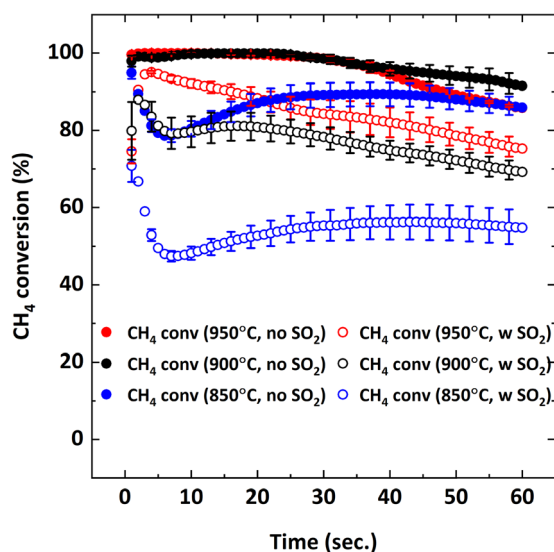


Fig. 11 CH<sub>4</sub> conversion as a function of reduction period during the 5th cycle with 9% CH<sub>4</sub>-balance N<sub>2</sub> and 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> at different temperatures (850, 900, 950 °C).

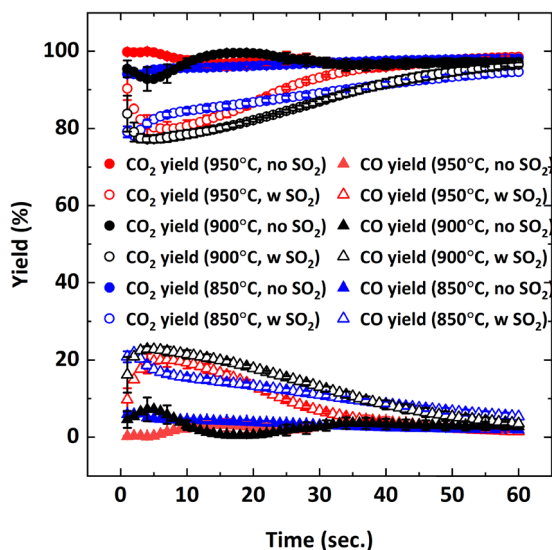


Fig. 12 CO<sub>2</sub> and CO yields as a function of reduction period during the 5th cycle with 9% CH<sub>4</sub>-balance N<sub>2</sub> and 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> at different temperatures (850, 900, 950 °C).

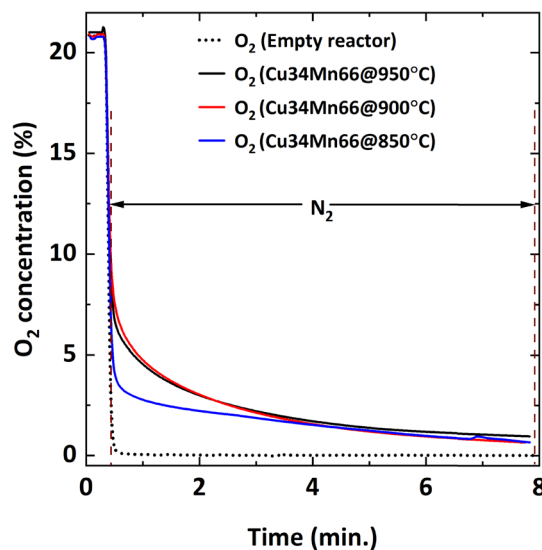


Fig. 13 O<sub>2</sub> concentration profile during oxygen uncoupling of Cu<sub>34</sub>Mn<sub>66</sub> particles at different temperatures (850, 900, 950 °C).

Effect of SO<sub>2</sub> on the combustion of CH<sub>4</sub> at different operating temperatures was also examined since equilibrium O<sub>2</sub> concentration and oxygen release rate of Cu<sub>34</sub>Mn<sub>66</sub> are directly proportional to temperature.<sup>36</sup> Fig. 11 illustrates CH<sub>4</sub> conversion and Fig. 12 illustrates the yields of CO<sub>2</sub> and CO at 850, 900 and 950 °C as a function of the reduction period with and without 5000 ppm SO<sub>2</sub> addition.

Fig. 11 shows that decreasing the operating temperature increases the negative effect of SO<sub>2</sub> on CH<sub>4</sub> conversion, with the effect being more pronounced at 850 °C. However, CO<sub>2</sub> and CO yields are similar at different temperatures (Fig. 12), indicating that SO<sub>2</sub> addition increases the unconverted CH<sub>4</sub>. To explain the effect of temperature, O<sub>2</sub> uncoupling profiles in N<sub>2</sub> at different temperatures are provided in Fig. 13. The O<sub>2</sub> concentration within the first 2 minutes is lower at 850 °C than those at 900 °C and 950 °C. That explains the lower CH<sub>4</sub> conversion at 850 °C than the conversion at 900 °C and 950 °C without SO<sub>2</sub>. When SO<sub>2</sub> is added, the decrease in CH<sub>4</sub> conversion is much larger comparing to the case without SO<sub>2</sub> since there is a competing O<sub>2</sub> demand between SO<sub>2</sub> and CH<sub>4</sub>. At 900 and 950 °C, O<sub>2</sub> amount is similar, but CH<sub>4</sub> conversion is higher at 950 °C in the presence of SO<sub>2</sub> while it is similar without SO<sub>2</sub>. Therefore, the negative effect of SO<sub>2</sub> on the reactivity of the oxygen carrier with CH<sub>4</sub> becomes more pronounced at lower temperatures.

#### Effect of SO<sub>2</sub> on multicycle stability of Cu<sub>34</sub>Mn<sub>66</sub>

The particles were exposed to 20 reduction-oxidation cycles with 1-minute reduction to determine if the CH<sub>4</sub> conversion degrades after multiple cycles when SO<sub>2</sub> is present. Fig. 14 shows that the concentration profiles of the flue gas species remain consistent throughout 20 cycles, indicating

that the performance of the particles does not degrade over multiple cycles. Following the exposure of oxygen carrier particles to SO<sub>2</sub> over 20 cycles, the particles were reduced without SO<sub>2</sub> with a mixture of 9% CH<sub>4</sub>-balance N<sub>2</sub>. The purpose of this experiment was to determine if the SO<sub>2</sub>-free reduction performance of the SO<sub>2</sub>-exposed particles differed from those not initially exposed to SO<sub>2</sub>. Fig. 14 shows that during the 21st cycle most of the inlet CH<sub>4</sub> converts to CO<sub>2</sub>, which is consistent with the behavior in Fig. 3. So, it can be assumed that the Cu<sub>34</sub>Mn<sub>66</sub> particles recover their reactivity after oxidation and after 20 cycles of SO<sub>2</sub> exposure no permanent deactivation occurs. Additionally, no agglomeration of the particles was observed after 21 cycles.

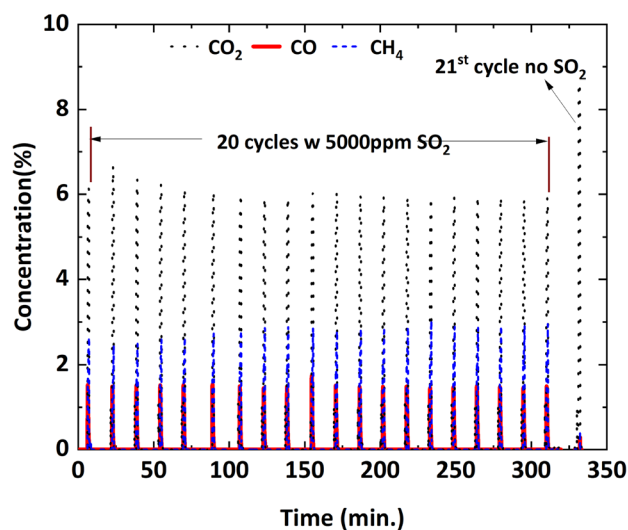


Fig. 14 Multicycle redox behavior of Cu<sub>34</sub>Mn<sub>66</sub> particles at 900 °C with 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub>. Reducing time is 1 minute.

### Characterization of Cu34Mn66

The Cu34Mn66 particles tested at various operating conditions were characterized with XRD, XPS, and TPR to evaluate any structural or chemical composition changes following SO<sub>2</sub> exposure. Characterization results of samples reduced with two different gas mixtures, *i.e.*, 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> and 9% CH<sub>4</sub>-balance N<sub>2</sub>, at 900 °C are reported here.

Fig. 15 shows the XRD patterns of Cu34Mn66 samples reduced for 1, 3 and 8 minutes for 5 cycles without SO<sub>2</sub> and the fresh oxidized sample for comparison. Powder XRD analysis confirms that fresh, fully oxidized Cu34Mn66 consists of CuMn<sub>2</sub>O<sub>4</sub> and Mn<sub>3</sub>O<sub>4</sub> crystalline structures. After 1-minute reduction, CuMn<sub>2</sub>O<sub>4</sub> converts into major Cu<sub>1.04</sub>Mn<sub>0.96</sub>O<sub>2</sub>, Mn<sub>3</sub>O<sub>4</sub>, and minor Cu phases in the absence of SO<sub>2</sub>. Extended reduction of Cu34Mn66 for 3 minutes results in additional Cu<sub>2</sub>O and MnO crystalline phases, while a fully reduced sample at 8 minutes consists of only Cu and MnO. Fig. 16 shows the XRD patterns for the samples reduced for 1, 3 and 8 minutes for 5 cycles in presence of 5000 ppm SO<sub>2</sub> and the oxidized sample following 1-minute reduction with SO<sub>2</sub>. Oxidized sample after 1-minute reduction with 5000 ppm SO<sub>2</sub> has an XRD pattern similar to the pattern of the fresh sample. Reduced samples exposed to SO<sub>2</sub> do not show any sulfate or sulfide presence and the XRD patterns look similar to the patterns with no SO<sub>2</sub> exposure. The only difference is for the sample reduced for 1 minute with SO<sub>2</sub> where the Cu peak is not present. Same thing is observed for the sample reduced for 1 minute for 20 cycles with and without SO<sub>2</sub> (Fig. S1 and S2 in ESI†). An interaction between elemental Cu and SO<sub>2</sub> likely occurs, but the resulting sulfur-copper compound might only be present on the surface and, therefore, not observed in the XRD pattern.

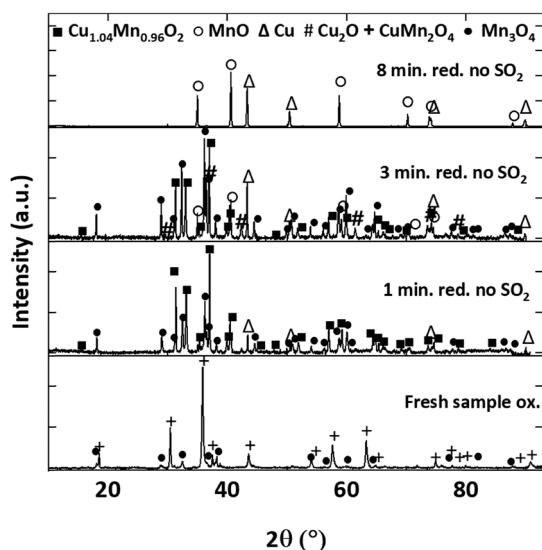


Fig. 15 XRD patterns of samples reduced for 1, 3 and 8 minutes for 5 cycles at 900 °C with 9% CH<sub>4</sub>-balance N<sub>2</sub> and a fresh oxidized sample.

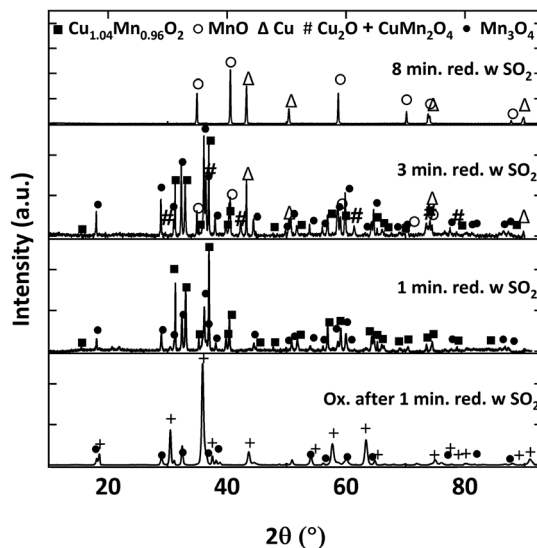


Fig. 16 XRD patterns of samples reduced for 1, 3 and 8 minutes for 5 cycles at 900 °C with 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub> and oxidized sample after 1-minute reduction for 5 cycles at 900 °C 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub>.

XPS analysis was performed for the samples exposed to SO<sub>2</sub> to determine if any sulfur species form on the surface of the oxygen carrier. Fig. 17 shows the S 2p spectra of 1, 3 and 8 minutes reduced samples with SO<sub>2</sub> exposure collected after the 1st cycle. According to literature, the S 2p<sub>3/2</sub> peak with a binding energy within 168.2–169.9 eV range indicates sulfate species presence on the surface.<sup>58–60</sup> S spectra is not observed for the 1-minute reduced sample. However, for 3 and 8 minutes reduced samples, sulfate presence on the surface of the reduced sample is observed. For the 8-minute reduction, there is an additional S 2p<sub>3/2</sub> position at a binding energy of 161.3 eV. This suggests the formation of sulfide species on a

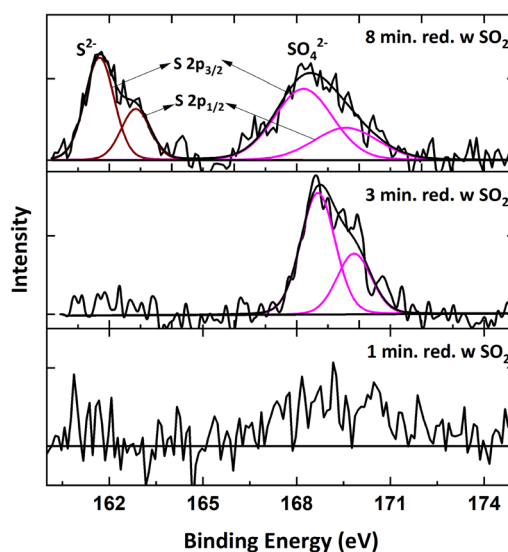


Fig. 17 S 2p spectra of samples reduced for 1, 3 and 8 minutes for 1 cycle at 900 °C with 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub>.

fully-reduced surface.<sup>58,59,61</sup> When the samples were exposed to SO<sub>2</sub> for multiple redox cycles, the sulfate peak is observed even for the 1-minute reduction after 5 cycles. The S spectra for 5th and 20th cycles are provided in Fig. S3†. For the 20th cycle, the sample was re-oxidized after being exposed to SO<sub>2</sub> for 20 cycles and the sulfate peak in the XPS spectra (Fig. S4†) is not distinctly seen after oxidizing the sample, which explains the behavior in Fig. 14 where the oxygen carrier is regenerated when it is oxidized and the effect of sulfur on the reactivity is not permanent.

Fig. 18 shows Cu 2p spectra of SO<sub>2</sub>-exposed samples after 1, 3 and 8 minutes of reduction after the 1st cycle. Deconvolution of spectra reveals three distinct Cu 2p<sub>3/2</sub> peaks at A (932.2–932.5 eV), B (933.8–934.3 eV), and C (935.7–936.1 eV) positions. On the other hand, in the absence of SO<sub>2</sub>, the reduced sample only has A and B peaks (Fig. S5†). From the literature, peak A corresponds to the Cu<sup>1+</sup> oxidation and may signify reduced CuMnO<sub>2</sub>, Cu, and Cu<sub>2</sub>O. Peak B indicates Cu<sup>2+</sup> oxides, such as CuO and CuMn<sub>2</sub>O<sub>4</sub>. Finally, peak C coincides with a Cu<sup>2+</sup> peak that ascribes to the Cu 2p<sub>3/2</sub> peak of copper sulfate compounds.<sup>59,62–66</sup> For the SO<sub>2</sub>-exposed samples reduced for 3 and 8 minutes, the presence of sulfate peak gets more prominent comparing to 1-minute reduction. For 1-minute reduction, the sulfate peak gets more pronounced as the number of cycles increases to 20 (Fig. S6†).

Fig. 19 shows Mn 2p spectra of reduced samples with SO<sub>2</sub> exposure after different reduction periods. From the literature, Mn 2p<sub>3/2</sub> peak positions at A (640.4–642.5 eV), B (641.3–641.9 eV), and C (641.6–646.2 eV) correspond to Mn<sup>2+</sup>, Mn<sup>3+</sup>, and Mn<sup>4+</sup> oxidation states, respectively.<sup>58,67–70</sup> Spectra for all three samples exhibit Mn<sup>2+</sup>, Mn<sup>3+</sup>, and Mn<sup>4+</sup> oxidation states. Since the Mn 2p<sub>3/2</sub> peak at 642.7 eV observed for MnSO<sub>4</sub> (ref. 71) overlaps with Mn<sup>4+</sup>, it is challenging to

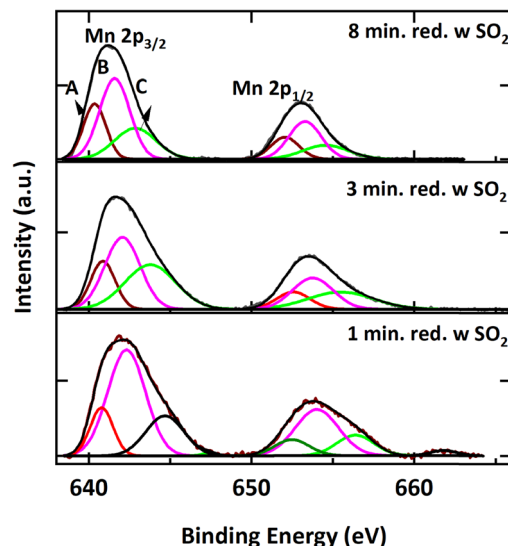


Fig. 19 Mn 2p spectra of samples reduced for 1, 3 and 8 minutes for 1 cycle at 900 °C with 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub>.

confirm the contribution of Mn towards sulfate formation on the surface.

To see if SO<sub>2</sub> is interacting with the Mn sites, TPR analysis was conducted. Fig. 20 shows the TPR profiles of three samples: fresh oxidized, 1-minute reduced sample with 5000 ppm SO<sub>2</sub> and 1-minute reduced sample without SO<sub>2</sub>. Each sample shows two broad overlapping peaks. One is a low-temperature peak between 305–325 °C, and the other is a high-temperature peak between 365–390 °C. The low and high-temperature peaks ascribe to the reduction of Cu<sup>2+</sup> to Cu, and the reduction of Mn<sup>3+</sup> and Mn<sup>4+</sup> to Mn<sup>2+</sup>, respectively.<sup>35,68,72</sup> The low-temperature peak shrinks and shifts to a lower temperature, while the high-temperature

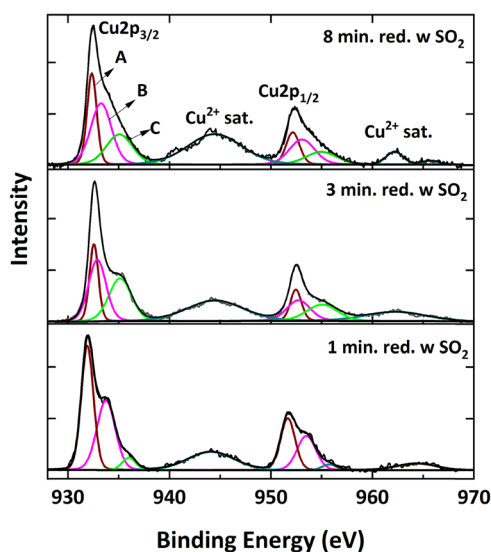


Fig. 18 Cu 2p spectra of samples reduced for 1, 3 and 8 minutes for 1 cycle at 900 °C with 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub>.

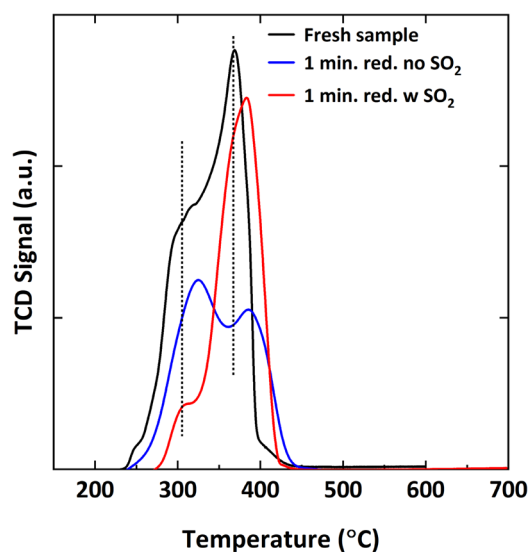


Fig. 20 H<sub>2</sub>-TPR spectra of fresh Cu<sub>34</sub>Mn<sub>66</sub> and samples reduced for 1 minute for 5 cycles at 900 °C with 9% CH<sub>4</sub>-balance N<sub>2</sub> and 9% CH<sub>4</sub>-5000 ppm SO<sub>2</sub>-balance N<sub>2</sub>.

peak gets more prominent when the sample is reduced in the presence of  $\text{SO}_2$  comparing to the sample reduced without  $\text{SO}_2$ . Since a clear difference in the reduction profiles for Mn is seen for the samples reduced with and without  $\text{SO}_2$ , it can be concluded that  $\text{SO}_2$  is interacting with Mn in addition to Cu.

## Discussion

The results of this study show that  $\text{SO}_2$  adversely affects  $\text{CH}_4$  conversion and, consequently, the performance of the Cu–Mn oxide in the CLOU process. The possible reasons for the adverse effects of  $\text{SO}_2$  are outlined in the following discussion.

According to previous studies, reactions between  $\text{CH}_4$  and  $\text{SO}_2$  in sub-stoichiometric conditions can form reduced  $\text{H}_2\text{S}$  and COS species<sup>73–75</sup> *via* heterogeneous catalytic reactions. Both Cu<sup>73</sup> and Mn<sup>76</sup> elements possess catalytic activity for these reactions. Similarly, they can also act as a catalyst for  $\text{SO}_2$  oxidation.<sup>77</sup> Experimental evidence in this study confirms  $\text{SO}_3$  presence, but there is no presence of reduced sulfur gaseous species such as  $\text{H}_2\text{S}$  or COS, not even during the extended reduction for 3 and 8 minutes.  $\text{SO}_2$  oxidation to  $\text{SO}_3$  explains the decrease in  $\text{O}_2$  concentration when  $\text{SO}_2$  is present during the reduction cycles when  $\text{CH}_4$  concentration is less than 9% (Fig. 10). When  $\text{SO}_2$  is added, the availability of  $\text{O}_2$  decreases because the excess  $\text{O}_2$  is being used up by  $\text{SO}_2$  to become oxidized to  $\text{SO}_3$ ; therefore, indicating a competing  $\text{O}_2$  demand between  $\text{SO}_2$  and  $\text{CH}_4$ .

Based on the XPS results,  $\text{SO}_3$  formation is not the only contributor in the reduction of  $\text{CH}_4$  conversion with the presence of  $\text{SO}_2$ . From the XPS results, S 2p and Cu 2p spectra confirm the presence of  $\text{SO}_4^{2-}$  species after 1-minute reduction, particularly in the form of  $\text{CuSO}_4$ . Sulfate species might form by the interaction of  $\text{SO}_2$  with any of the reduced species such as Cu,  $\text{CuMnO}_2$ , or  $\text{Mn}_3\text{O}_4$ . The presence of manganese-sulfur surface species remains unconfirmed *via* XPS. Still, the formation of Mn-sulfate species on the surface of the oxygen carrier is possible since previous selective catalytic reduction (SCR) studies show that  $\text{SO}_2$  could poison Mn-based catalysts.<sup>68</sup> In fact, the TPR results shows that  $\text{SO}_2$  not only interacts with Cu, but also Mn sites. On the other hand, XRD analysis reveals no sulfate or sulfide crystalline structure formation despite the presence of sulfate on the surface of the oxygen carrier. Therefore, it can be concluded that the crystalline structure of the oxygen carrier is not affected by  $\text{SO}_2$ .

Decreased  $\text{CH}_4$  conversion in the presence of  $\text{SO}_2$  could be as a result of the sulfate species on the surface inhibiting  $\text{O}_2$  release and/or the heterogeneous CLC reaction *via* lattice oxygen. However, no change in  $\text{O}_2$  release was observed in this study (Fig. 2). Therefore, the reduction in  $\text{CH}_4$  conversion could be attributed to the heterogeneous CLC reaction being impacted by the presence of sulfate species on the surface of the oxygen carrier, in addition to the

competing oxygen demand between  $\text{SO}_2$  and  $\text{CH}_4$  using the released  $\text{O}_2$  in the gas phase.

Despite the observed decrease in activity, the effect of  $\text{SO}_2$  does not worsen over multiple redox cycles. Moreover, the  $\text{SO}_2$ -exposed sample can be regenerated after oxidation, going back to its original activity before  $\text{SO}_2$  exposure.

The extent of reduction of Cu–Mn oxide also affects the influence  $\text{SO}_2$  has on the oxygen carrier and, consequently, the fuel conversion. A high degree of reduction produces significant elemental Cu, increasing  $\text{SO}_2$  interaction probability, as found during the 3- and 8-minute reduction cycles. For instance, as seen from the XPS analysis, sulfide species form on the surface of the oxygen carrier upon full reduction. Keeping the reduction extent of materials within the CLOU mechanism with a short residence time where there is less Cu might improve the performance of the Cu–Mn oxide. In addition, due to  $\text{SO}_3$  formation, the oxygen carrier to fuel ratio might need to be increased to compensate for the oxygen needed for  $\text{SO}_3$  formation. It is possible that with a higher oxygen carrier to fuel ratio, the adverse effect of sulfur may not be significant.

## Conclusions

The interaction of  $\text{SO}_2$  with a bimetallic Cu–Mn oxygen carrier has been investigated under CLOU conditions in a batch fluidized-bed reactor. The results suggest that  $\text{SO}_2$  negatively affects total  $\text{CH}_4$  conversion and  $\text{CO}_2$  yield. However, the oxygen release capacity of the Cu–Mn oxide remains unaffected. The reactivity with  $\text{CH}_4$  decreases as the  $\text{SO}_2$  concentration increases. Increasing the temperature from 850 to 950 °C improves the conversion of  $\text{CH}_4$ . During the reduction cycle,  $\text{SO}_2$  oxidizes to  $\text{SO}_3$ , creating a competing oxygen demand in the reactor with the fuel, and lowers  $\text{CH}_4$  conversion if there is not enough oxygen for both reactions. Hence, decreasing the feed  $\text{CH}_4$  concentration with the same oxygen carrier amount lowers the impact of  $\text{SO}_2$  on  $\text{CH}_4$  conversion.

Additionally, characterization of  $\text{SO}_2$ -exposed oxygen carriers *via* XPS shows that sulfate species form on the surface, possibly causing further reduction in  $\text{CH}_4$  conversion. However, the presence of sulfate does not cause a permanent deactivation of the particles. The reduction in  $\text{CH}_4$  conversion does not worsen during 20 redox cycles with 5000 ppm  $\text{SO}_2$  and the particles can be fully regenerated during the oxidation cycle. Nevertheless, further studies are required in a continuous CLOU unit to determine the sulfur effect during a long-time operation.

## Author contributions

Turna Barua: investigation, methodology, formal analysis, writing – original draft. Bihter Padak: conceptualization, funding acquisition, methodology, project administration, supervision, writing – review & editing.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

The authors are grateful for the financial support provided by the National Science Foundation's (NSF) Combustion and Fire Systems Program (award number 1606874). The authors would also like to thank Madeline Talebi for her assistance with conducting some of the experiments and synthesis of some of the oxygen carriers used in the experiments. She has also helped editing the language of the paper.

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