

PAPER

[View Article Online](#)
[View Journal](#) | [View Issue](#)


Cite this: *Green Chem.*, 2020, **22**, 5059

A highly active Pt/In₂O₃ catalyst for CO₂ hydrogenation to methanol with enhanced stability†

Kaihang Sun,^{a,b} Ning Rui,^{a,b} Zhitao Zhang,^{a,b} Zeyu Sun,^{a,b} Qingfeng Ge^c and Chang-Jun Liu  ^{a,b}

Supported Pt catalysts have been extensively investigated for CO₂ hydrogenation with methane and CO as the principal products. In this work, a Pt/In₂O₃ catalyst prepared with decomposition precipitation was tested for CO₂ hydrogenation to methanol. The Pt/In₂O₃ catalyst exhibited a highly improved activity towards CO₂ hydrogenation, with methanol selectivity of ca. 100% at temperatures below 225 °C, 74% at 275 °C and 54% at 300 °C, respectively, compared to the pure In₂O₃ catalyst. This represents the highest methanol selectivity reported on Pt catalysts for CO₂ hydrogenation. The stability of the Pt/In₂O₃ catalyst at elevated temperatures has also been higher than that of the pure In₂O₃ catalyst, indicated by the methanol formation rate decreasing only to 95% of the initial rate after 5 h in the reaction stream and remaining largely constant thereafter. In contrast, the pure In₂O₃ catalyst loses 20% of the initial methanol formation rate after 9 h in the reaction stream. The characterization of the catalysts confirms that the Pt nanoparticles are well dispersed on In₂O₃ with a particle size below 3 nm. The strong metal–support interaction (SMSI) between Pt and In₂O₃ improves the stability of the catalyst and prevents the over-reduction of In₂O₃. The synergy between the supported Pt nanoparticles and In₂O₃ balances the hydrogen activation and the density of the surface oxygen vacancies in In₂O₃, resulting in the high activity for CO₂ hydrogenation and enhanced stability of the Pt/In₂O₃ catalyst.

Received 9th May 2020,
Accepted 26th June 2020

DOI: 10.1039/d0gc01597k

rsc.li/greenchem

Introduction

With the development of renewable energy, CO₂ capture and utilization^{1–4} have the potential to alleviate the elevated CO₂ level. Hydrogenation of CO₂ to methanol⁵ has been considered as key to the so-called “methanol economy”.⁶ Meanwhile, this process can also be considered as a means of hydrogen storage and/or energy storage.⁷ Extensive and continual efforts have been made over the past decades to develop a highly active catalyst with high methanol selectivity. The investigated catalysts include those based on Cu/ZnO,^{8–10} ZnO–ZrO₂,¹¹ GaNi alloy,¹² bimetallic PdZn¹³ and others. Our group predicted *via* density functional theoretical (DFT) studies, and then confirmed experimentally, that In₂O₃ with oxygen vacancies exhibited a high activity for CO₂ activation and further high selectivity for hydrogenation to methanol.^{14–18} Since then,

In₂O₃-based catalysts have attracted increasing attention as novel catalysts for CO₂ hydrogenation to methanol. Further studies include novel preparation and characterization of In₂O₃,^{19–21} combination of In₂O₃ with other oxides, such as ZrO₂,^{22–25} and the addition of metals, including copper,^{26,27} cobalt,²⁸ palladium,^{18,29} rhodium,³⁰ and nickel,³¹ to further improve the catalytic activity for CO₂ activation and hydrogenation. It has also been demonstrated that the strong hydrogenation ability of the metal catalysts enhances the hydrogenation of the In₂O₃-supported metal catalysts.^{18,26–33} Among the metals considered, platinum is an important one for its extensive applications in hydrogenation and photo-catalytic conversion.³³ Although a Pt/In₂O₃ film catalyst has been shown to exhibit an enhanced activity for CO₂ hydrogenation to methanol, the reaction was performed under atmospheric pressure plasma conditions.³² The In₂O₃-supported Pt catalyst has not been examined for thermal catalytic methanol synthesis from CO₂ and hydrogen.

The supported Pt catalysts normally exhibit a high activity for CO₂ hydrogenation to methane and CO. The metal dispersion,^{34–36} the acidity of the support,³⁵ the effect of the promoter³⁷ and oxygen vacancies,³⁶ the metal–support interaction,^{34,35,38,39} and the degree of Pt reduction^{40,41} have been shown to affect the activity and selectivity. CO₂ hydrogenation to methanol has been investigated as well.^{42,43} But the

^aTianjin Co-Innovation Center of Chemical Science & Engineering, Tianjin University, Tianjin 300072, China. E-mail: ughg_cjl@yahoo.com

^bSchool of Chemical Engineering and Technology, Tianjin University, Tianjin 300072, China

^cDepartment of Chemistry and Biochemistry, Southern Illinois University, Carbondale, Illinois 62901, USA

†Electronic supplementary information (ESI) available. See DOI: 10.1039/d0gc01597k

most selective products remain to be methane and CO over the Pt catalysts.^{34–36} A Pt catalyst with high methanol selectivity from CO₂ hydrogenation is desired.

Although In₂O₃-based catalysts are active and selective for methanol synthesis from CO₂ and H₂, the over-reduction of In₂O₃ under the reaction conditions causes deactivation and a decrease in selectivity because In₂O₃ can be reduced readily in the presence of hydrogen and CO.^{19,44,45} Over-reduced In₂O₃ catalyses preferably the competing reverse water gas shift (RWGS) reaction, resulting in CO instead of methanol, especially at elevated reaction temperatures. CO is more potent at reducing In₂O₃ than H₂. Its production during the reaction exacerbates the problem, leading to further reduction of the activity and selectivity.¹⁹

In this work, we prepared the Pt/In₂O₃ catalysts using the decomposition precipitation method. The resulting catalyst shows an improved activity over In₂O₃ alone. High methanol selectivity was achieved on the Pt/In₂O₃ catalyst with a methanol selectivity of *ca.* 100% at temperatures below 225 °C. The methanol selectivity was maintained at 74% at 275 °C and 54% at 300 °C. Inhibition of the over-reduction of In₂O₃ was achieved due to the strong metal-support interaction (SMSI) between Pt and In₂O₃, resulting in the enhanced stability.

Experimental

Catalyst preparation

In₂O₃ was prepared using the precipitation method. In (NO₃)₃·4H₂O (HWRK Chem, 99.99%) and Na₂CO₃·10H₂O (Tianjin Kermel Chemical Reagent, 99%) are the precursor and precipitant, respectively. Firstly, a desired amount of the precursor was dissolved in 100 mL of deionized water (0.15 mol L⁻¹), followed by vigorous stirring at 70 °C until fully dissolved. Then 0.2 mol L⁻¹ of Na₂CO₃ was added to the solution dropwise at the same temperature until the pH value reached 7. The mixture was aged at 80 °C for 2 h after being stirred for an additional 0.5 h. The precipitate was then washed with deionized water several times. The resulting solid was dried at 80 °C for 12 h prior to calcination in static air at 450 °C for 3 h.

The Pt/In₂O₃ catalyst was prepared using the deposition-precipitation method. Platinum nitrate hydrate (Aladdin Chemicals, Shanghai, 18.02 wt% of metal platinum basis) was first dissolved in 50 mL of deionized water. The as-prepared In₂O₃ was then added into the solution, and the resulting mixture was stirred vigorously for 1 h at room temperature. This was followed by adding an excess amount of urea (0.2 g) into the mixture under continuous stirring at 80 °C for 3 h while the precipitate formed. The precipitate was washed and filtered with 500 mL of deionized water. Prior to calcination in static air at 450 °C for 3 h, the precipitate was freeze-dried overnight.

Characterization

The catalyst samples were characterized using X-Ray fluorescence (XRF) spectroscopy on a JEOL Supermini200 system,

equipped with a palladium X-ray source (Pd anode, 50 kV and 200 W). 77 K N₂ adsorption/desorption isotherms of the samples were measured on an AUTOSORB-1-C instrument (Quantachrome). The specific surface area (*S*_{BET}) was calculated using the Brunauer-Emmett-Teller (BET) model.

The crystal structure of the catalyst was determined by powder X-ray diffraction (PXRD), using a Rigaku D/max 2500v/pc diffractometer with Cu K α radiation (40 kV, 200 mA) at a scanning rate of 1° min⁻¹ within the 2 θ range of 10°–90°. The phase identification was performed by comparing with the Joint Committee on Powder Diffraction Standards (JCPDS).

Transmission electron microscopy (TEM) images were collected on a JEOL JEM-2100F system equipped with an energy-dispersive X-ray spectrometer (EDX) operated at 200 kV. The sample powder was suspended in ethanol and dispersed ultrasonically for 10 min. A drop of the suspension was deposited on a copper grid coated with carbon.

Temperature-programmed reduction of hydrogen (H₂-TPR) was conducted on a Micromeritics Autochem II 2920 chemisorption analyzer. About 200 mg of the sample was placed into a U-shaped quartz tube. The tube with the sample was purged by flowing helium for 1 h at 200 °C to remove the absorbed impurities. After cooling to –50 °C with liquid nitrogen under flowing helium, the sample was heated to 350 °C at a rate of 10 °C min⁻¹ under a gaseous mixture of 10% H₂ in N₂ as the reductant. The effluent gas was analysed with a TCD. The TCD signal was normalized to the weight of the sample for subsequent analysis.

Temperature-programmed desorption of carbon dioxide (CO₂-TPD) was conducted on the Micromeritics Autochem II 2920 chemisorption analyzer equipped with a Hiden HPR-20 mass spectrometer (MS). About 100 mg of the sample was placed in a U-shaped quartz tube and reduced with 10% H₂/Ar for 1 h at 200 °C and then cooled to 50 °C under flowing helium, followed by CO₂ adsorption at the same temperature for 1 h. The tube with the sample was purged with flowing helium for 1 h to remove physically adsorbed CO₂. It was then heated to 500 °C at a rate of 10 °C min⁻¹ under flowing helium. The effluent gas was analysed using the TCD and mass spectrometer (the signal of *m/z* = 44 was collected from the TCD outlet). The signal was also normalized to the weight of the sample for subsequent analysis.

Raman spectra were acquired using an inVia Reflex Renishaw Raman Spectroscopy System in the range of 200–800 nm, with a 532 nm laser as the excitation source. The laser power was set at 7 mW and the integration time was 5 s. The Raman signals were normalized to the weight of the sample for the subsequent analysis.

The UV-vis absorption spectra of the samples were recorded with a UV-2600 UV-vis spectrophotometer (Shimadzu Corporation) to investigate the formation of surface oxygen vacancies.

Catalytic activity test

The catalytic activity for CO₂ hydrogenation to methanol over the Pt/In₂O₃ catalyst and the pure In₂O₃ catalyst was tested in a

vertical fixed bed reactor. 0.2 g of catalyst was diluted with 1.0 g of SiC before being loaded into the reactor. Prior to the test, the reactor with the catalyst was purged with N₂ (30 ml min⁻¹) for 0.5 h at room temperature. The Pt/In₂O₃ catalyst was pre-reduced under flowing hydrogen (H₂/N₂ = 1/9, molar ratio, 30 ml min⁻¹) at 200 °C for 1 h. The reactant mixture (H₂/CO₂/N₂ = 76/19/5, molar ratio) was then introduced into the reactor until the pressure reached 5 MPa. The catalytic activity was measured from 200 °C to 300 °C at a constant flow rate of 70 ml min⁻¹. The gaseous hourly space velocity (GHSV) was adjusted to 21 000 cm³ h⁻¹ g_{cat}⁻¹. The effluent was analysed with an online gas chromatograph (Agilent 7890A) equipped with two columns, one connected to a flame ionized detector (FID) and the other to the thermal conductivity detector (TCD). The valves and lines between the reactor outlet and the GC inlet were maintained at 110 °C to prevent the condensation of methanol.

The CO₂ conversion (X_{CO_2}), methanol selectivity (S_{methanol}), and methanol formation rate were calculated according to the following equations:

$$X_{\text{CO}_2} = \frac{F_{\text{CO}_2, \text{in}} - F_{\text{CO}_2, \text{out}}}{F_{\text{CO}_2, \text{in}}} \times 100\%,$$

$$S_{\text{methanol}} = \frac{F_{\text{methanol, out}}}{F_{\text{CO}_2, \text{in}} - F_{\text{CO}_2, \text{out}}} \times 100\%,$$

$$\text{Methanol formation rate} = \frac{F_{\text{CO}_2, \text{in}} \times X_{\text{CO}_2} \times S_{\text{methanol}}}{W} \times M,$$

where F is the molar flow rate, M the molar mass of methanol and W the weight of the catalyst.

Results and discussion

Catalytic activity tests

Fig. 1a shows CO₂ conversion and methanol selectivity on the pure In₂O₃ and Pt/In₂O₃ catalysts. Obviously, the conversion of CO₂ increases when the reaction temperature is increased. The methanol selectivity is close to 100% at reaction temperatures below 225 °C. It decreases when the temperature is increased further as the competing RWGS reaction is endothermic. On the other hand, the selectivity toward methanol is still higher than 70% at 275 °C and 50% at 300 °C. These results indicate that the Pt/In₂O₃ catalyst remains highly selective for methanol synthesis from CO₂ hydrogenation. Furthermore, as shown in Fig. 1b, the Pt/In₂O₃ catalyst shows a significantly higher methanol formation rate than the pure In₂O₃ catalyst at the same temperature. The CO₂ conversion reaches 17.3% at 300 °C for the Pt/In₂O₃ catalyst with a methanol formation rate of 0.542 g_{methanol} h⁻¹ g_{cat}⁻¹, significantly higher than 0.335 g_{methanol} h⁻¹ g_{cat}⁻¹ on the pure In₂O₃ catalyst with a CO₂ conversion of 9.4%. These results clearly demonstrate that the

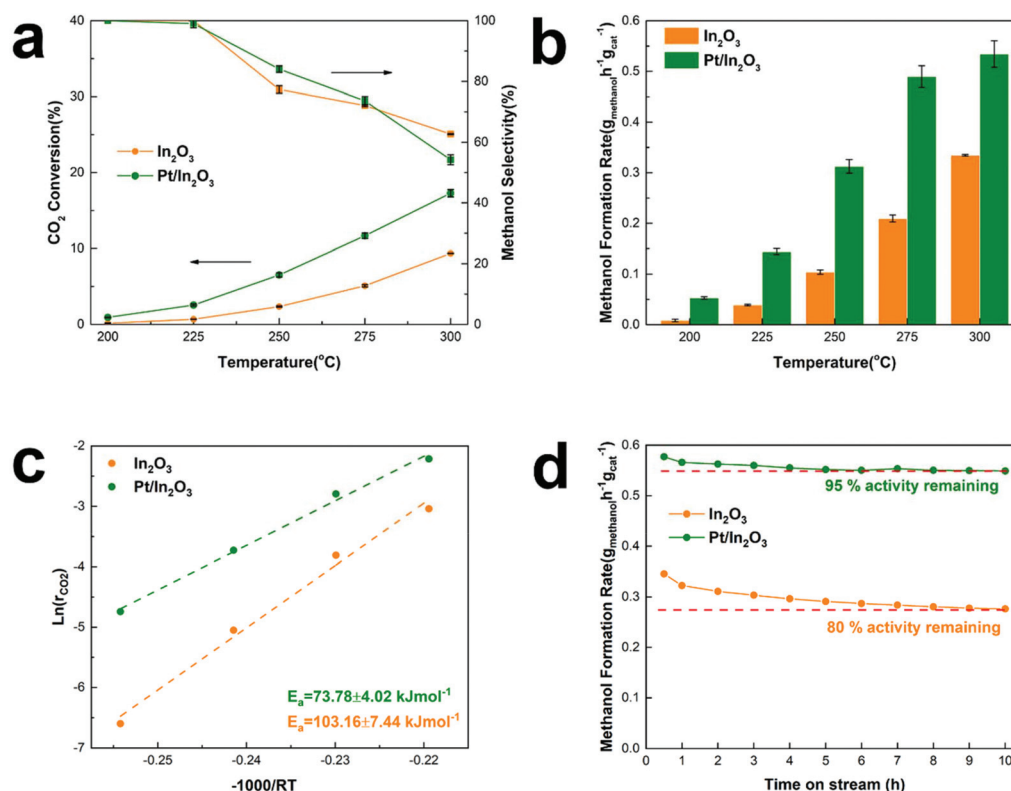


Fig. 1 Catalytic activity tests of the Pt/In₂O₃ and pure In₂O₃ catalysts. (a) CO₂ conversion and methanol selectivity, (b) methanol formation rate as a function of reaction temperature, (c) apparent activation energy of CO₂ conversion and (d) methanol formation rate versus time of the catalyst in the reaction stream (TOS).

Pt loading (1.07 wt% according to the results of XRF) on In_2O_3 significantly increases the CO_2 conversion and the methanol formation rate. The comparison of typically reported catalysts for CO_2 hydrogenation to methanol is shown in Table S1.† The $\text{Pt}/\text{In}_2\text{O}_3$ catalyst shows a higher activity than most of the reported catalysts. Compared to the supported Ni and Rh catalysts in Table S1,† the $\text{Pt}/\text{In}_2\text{O}_3$ catalyst shows better CO_2 conversion with higher selectivity towards methanol at lower temperatures of 250 °C and 275 °C.

The methanol selectivity of the $\text{Pt}/\text{In}_2\text{O}_3$ catalyst at 300 °C is slightly lower than that of the pure In_2O_3 catalyst at the same temperature, due largely to the agglomeration of the Pt nanoparticles at elevated temperatures. The average size of the Pt nanoparticles increases from 1.47 nm to 2.42 nm after reaction at 300 °C and 5 MPa. The large-sized Pt nanoparticles facilitate RWGS reaction according to the literature.²⁹ We speculate that hydrogenation of CO_2 over Pt nanoparticles prefers to undergo the RWGS + CO-hydrogenation pathway. Weakening the binding of CO on the Pt nanoparticles leads to higher selectivity to CO.³⁸ Thus, the methanol selectivity of the $\text{Pt}/\text{In}_2\text{O}_3$ catalyst is lower than that over the pure In_2O_3 catalyst at 300 °C.

The apparent activation energy of the pure In_2O_3 and the $\text{Pt}/\text{In}_2\text{O}_3$ catalysts was calculated based on the Arrhenius equation at a CO_2 conversion of less than 15% to minimize the effect of mass and heat transfer. As shown in Fig. 1c, the activation energy of CO_2 conversion over the $\text{Pt}/\text{In}_2\text{O}_3$ catalyst is 73.78 kJ mol^{-1} , much lower than that over the pure In_2O_3 catalyst (103.16 kJ mol^{-1}). This indicates that the Pt loading on In_2O_3 significantly reduces the activation barrier for CO_2 conversion. The apparent activation energy for methanol and CO production was calculated and is shown in Fig. S1.† For methanol production, the activation energies on the $\text{Pt}/\text{In}_2\text{O}_3$ and pure In_2O_3 catalysts are 64.08 kJ mol^{-1} and 91.83 kJ mol^{-1} , respectively. This indicates that the supported Pt nanoparticles reduce the activation energy of methanol production

significantly from that on the pure In_2O_3 catalyst. In contrast, the activation energies for CO production are 102.18 kJ mol^{-1} and 96.48 kJ mol^{-1} on the $\text{Pt}/\text{In}_2\text{O}_3$ and pure In_2O_3 catalysts, respectively. These results demonstrate that supported Pt over In_2O_3 promotes methanol production, while it suppresses the RWGS reaction.

Tsoukalou *et al.*¹⁹ reported that In_2O_3 was not stable over the course of the CO_2 hydrogenation reaction and could not maintain its initial activity with a 200 min test. They concluded that the formation of In^0 as a result of partial reduction of In_2O_3 caused deactivation,¹⁹ which typically occurs at elevated temperatures. The present results show that the stability of the pure In_2O_3 catalyst can be significantly improved in the presence of well-dispersed Pt nanoparticles. As shown in Fig. 1d, the methanol formation rate on the $\text{Pt}/\text{In}_2\text{O}_3$ catalyst only decreased slightly from its initial value and was stabilized at 95% of the initial rate after a 5 h reaction at 300 °C and 5 MPa. In contrast, the pure In_2O_3 catalyst needs 9 h to become stable under the same reaction conditions with a much more significant reduction (20%) in the methanol formation rate.

X-ray diffraction and TEM analysis

Based on the results of N_2 adsorption, the specific surface areas of the $\text{Pt}/\text{In}_2\text{O}_3$ and pure In_2O_3 catalysts were determined to be 63.8 $\text{m}^2 \text{g}^{-1}$ and 65.8 $\text{m}^2 \text{g}^{-1}$, respectively. The XRD patterns of the $\text{Pt}/\text{In}_2\text{O}_3$ and pure In_2O_3 catalysts are shown in Fig. 2. The $\text{Pt}/\text{In}_2\text{O}_3$ catalyst before hydrogen reduction was named “ $\text{Pt}/\text{In}_2\text{O}_3\text{-BT}$ ”. The samples after reaction at 300 °C and 5 MPa were labelled as “ $\text{Pt}/\text{In}_2\text{O}_3\text{-AR}$ ” and “ $\text{In}_2\text{O}_3\text{-AR}$ ”. The diffraction peaks at 21.5°, 30.7°, 35.5°, 45.7°, 51.0° and 60.7° were assigned to the diffraction from the (211), (222), (400), (431), (440) and (622) planes of the $Ia\bar{3}$ (206) space group according to PDF#06-0416. No characteristic diffraction peaks of the Pt species could be observed in the XRD patterns, indicating that Pt is highly dispersed on In_2O_3 with sizes below 3 nm. Fig. 2 also indicates that the crystal structure of the $\text{Pt}/\text{In}_2\text{O}_3$ catalyst remains intact after hydrogen reduction and reaction.

Figs. 3a and b show the TEM images of the $\text{Pt}/\text{In}_2\text{O}_3$ catalyst. As shown in Fig. 3c, the well-dispersed Pt nanoparticles, with an average diameter of 1.47 nm, on In_2O_3 can be clearly observed. The lattice spacing of the Pt particle on In_2O_3 is 0.2272 nm, which can be attributed to the (111) planes. As shown in Figs. 3d and e, the Pt nanoparticles exposing the (111) facet with a lattice spacing of 0.2310 nm can also be observed after reaction at 300 °C and 5 MPa. The increase can be attributed to the diffusion of Pt nanoparticles towards the bulk region of In_2O_3 .²⁹ The average diameter increases to 2.42 nm but is still below 3 nm as shown in Fig. 3f. The samples annealed at high temperature (>700 K) appear to undergo a classical SMSI transition, as has been observed for Pt and TiO_2 .⁴⁶ Therefore, a thin In_2O_3 overlayer covering the Pt nanoparticles indicates that the SMSI between Pt and In_2O_3 exists before and after the reaction.^{47–49}

Figs. 4a and b show the TEM images of the pure In_2O_3 catalyst before the reaction and Figs. 4c and d show those after the

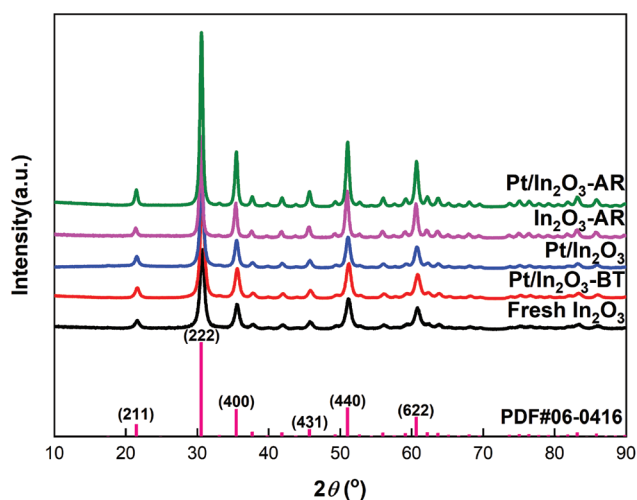


Fig. 2 XRD patterns of the $\text{Pt}/\text{In}_2\text{O}_3$ and pure In_2O_3 catalysts.

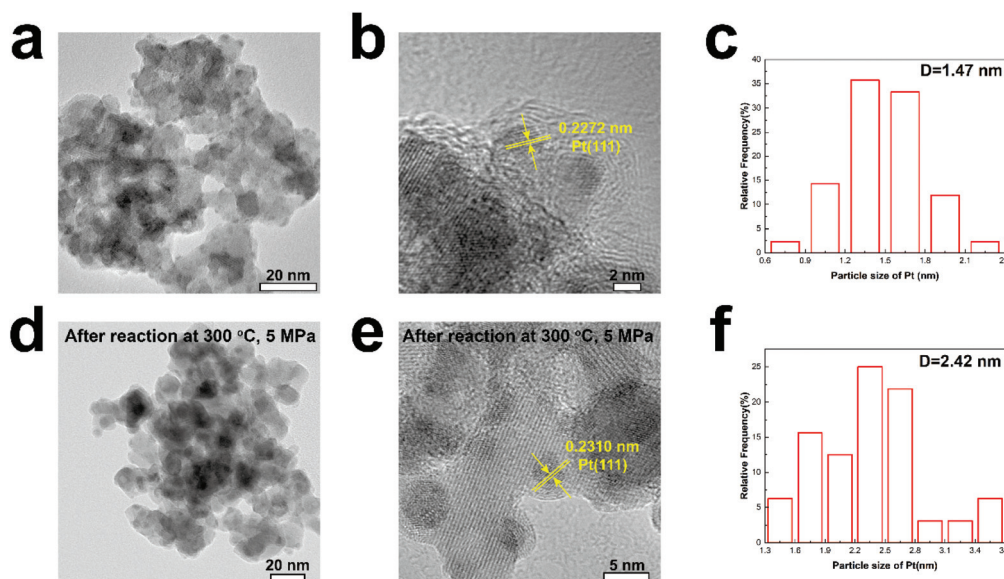


Fig. 3 (a) and (b) TEM images of the Pt/In₂O₃ catalyst, (c) particle size distribution for the Pt/In₂O₃ catalyst, (d) and (e) TEM images of Pt/In₂O₃ after reaction, and (f) particle size distribution for Pt/In₂O₃ after reaction.

reaction. The In₂O₃ particles can be clearly identified with lattice fringes attributable to the (222) and (211) planes, suggesting that the catalyst exists in the form of In₂O₃ and retains good crystallinity, consistent with the XRD results.

Raman spectroscopy

Raman spectroscopy has been believed to provide direct spectroscopic evidence of the oxygen vacancies in In₂O₃.^{50,51} In the present study, the catalysts were analysed using room temperature-visible Raman scattering to probe the oxygen vacancies in

In₂O₃. The Raman spectra are shown in Fig. 5a. The peak at 131 cm⁻¹ is assigned to the In-O vibration of the InO₆ structural units.⁵² The peak at 307 cm⁻¹ is usually interpreted as the bending vibration $\delta(\text{InO}_6)$ of octahedra.⁵³ The two peaks at 497 cm⁻¹ and 629 cm⁻¹ are attributed to the stretching vibration of the same $\nu(\text{InO}_6)$ octahedra.⁵⁰ The scattering feature at 367 cm⁻¹ can be attributed to the stretching vibration of In-O-In ($\nu(\text{In-O-In})$) and it reflects the presence of oxygen vacancies in the In₂O₃.^{15,50,51,54,55} The intensities of the 367 cm⁻¹ peak of Pt/In₂O₃ and In₂O₃-R (In₂O₃ after hydrogen reduction) increase significantly compared to those of the fresh samples, consistent with the expectation that hydrogen reduction creates more oxygen vacancies.

In order to quantify the oxygen vacancies, we integrated the peaks at 307 cm⁻¹ (*I*₁) and 367 cm⁻¹ (*I*₂) and used the *I*₂/*I*₁ ratio to characterize the amount of the oxygen vacancies on In₂O₃.⁵⁶ As shown in Fig. 5b, the oxygen vacancies in the Pt/In₂O₃-BT sample are 81% of those of the pure In₂O₃ catalyst, likely due to the presence of the PtO_x species covering some oxygen vacancies. However, the amount of the oxygen vacancies on the Pt/In₂O₃ catalyst increases to 92% of the In₂O₃-R sample after hydrogen reduction. This suggests that the well-dispersed Pt nanoparticles have a favourable effect on the formation of the oxygen vacancies over In₂O₃.

More importantly, the amount of the oxygen vacancies on the pure In₂O₃ and the Pt/In₂O₃ catalysts varies quite differently after the reaction at 300 °C and 5 MPa. In₂O₃-AR has many more oxygen vacancies than Pt/In₂O₃-AR. However, the catalytic activity and stability of the pure In₂O₃ catalyst are worse than those of the Pt/In₂O₃ catalyst. These results demonstrate that the additional oxygen vacancies created at high temperature in In₂O₃ had a negative effect on the reaction, causing In₂O₃ over-reduction and destabilization of the cata-

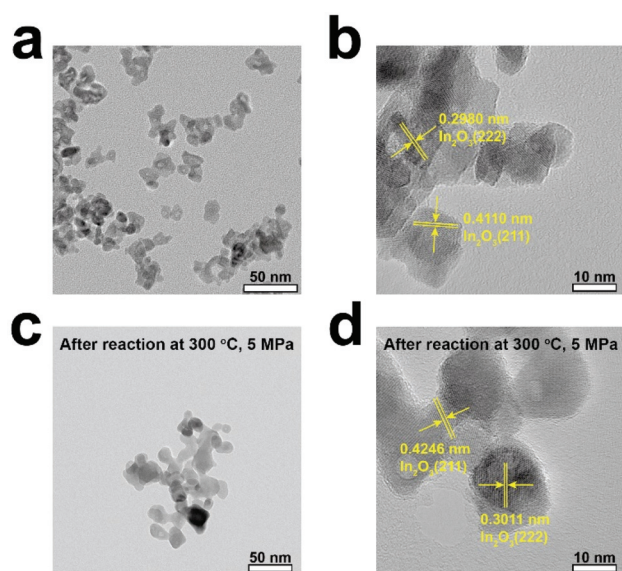


Fig. 4 TEM images of (a and b) the fresh In₂O₃ catalyst and (c and d) the pure In₂O₃ catalyst after reaction.

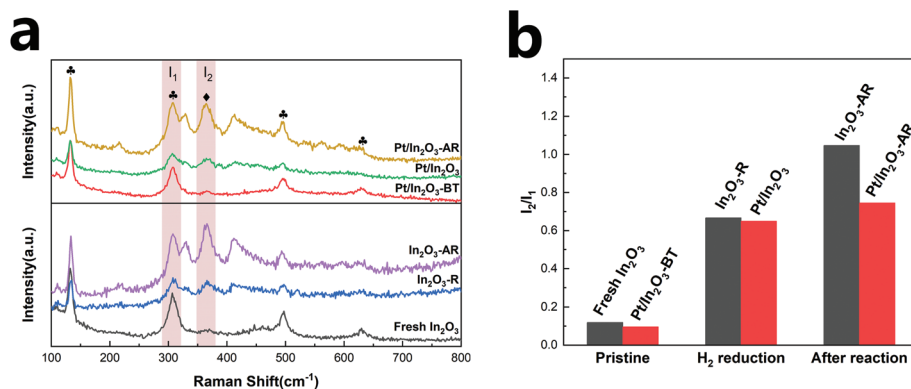


Fig. 5 (a) Visible Raman spectra of the Pt/In₂O₃ and pure In₂O₃ catalysts, (b) comparison of the relative amount of the oxygen vacancies according to the I₂/I₁ ratio. ♣ : InO₆ octahedron, ♦ : the stretching vibrations of In-O-In. In₂O₃-R represents the In₂O₃ catalyst after hydrogen reduction (H₂/N₂ = 1/9, molar ratio, at 200 °C for 1 h).

lyst. This is, in general, consistent with Tsoukalou *et al.*,¹⁹ who showed that the over-reduction of In₂O₃ results in the formation of metallic indium, which is inactive for hydrogen activation.

Hydrogen temperature programmed reduction

H₂-TPR was used to characterize the reduction behaviour of the catalyst and the formation of oxygen vacancies. As shown in Fig. 6a, the reduction peak at 133 °C of pure In₂O₃ can be assigned to the reduction of the surface, corresponding to the formation of the surface oxygen vacancies.^{15,18,54} The reduction peaks appear at much lower temperatures of 11 °C and 41 °C for the Pt/In₂O₃-BT sample. These peaks can be assigned to the reduction of the PtO_x species with the assistance of the Pt nanoparticles.²⁹ The formation of PtO_x also indicates the existence of the SMSI between Pt and In₂O₃, which in turn promotes dispersion of the Pt nanoparticles.^{47,57,58} Furthermore, the active H adatoms spilled over from the Pt

nanoparticles to In₂O₃ can also cause the reduction of the In₂O₃ surface,^{18,29,59} making the In₂O₃ surface reduction peak disappear. The normalized intensity of the reduction peak of the Pt/In₂O₃-BT sample is much higher than that of In₂O₃. This also confirms that the well-dispersed Pt nanoparticles promote the formation of the surface oxygen vacancies of In₂O₃ during hydrogen reduction. This is in line with the results of the Raman analyses discussed above.

CO₂ temperature programmed desorption

Generally, the surface oxygen vacancies on In₂O₃ promote CO₂ adsorption and activation.¹⁵ Therefore, CO₂ can be used as a probe molecule to characterize the surface oxygen vacancies on In₂O₃ through CO₂-TPD.⁶⁰ The profiles of CO₂-TPD in Fig. 6b confirm that less CO₂ is adsorbed on the Pt/In₂O₃ catalyst than on the pure In₂O₃ catalyst. The peak located at *ca.* 100 °C for these two samples is assigned to the physically adsorbed CO₂. It is noteworthy that the normalized intensity of the desorption peak at 357 °C for fresh In₂O₃ is much stronger than that for the Pt/In₂O₃ catalyst. This peak can be attributed to CO₂ adsorption in the hydrogen-induced surface oxygen vacancies.¹⁸ The peaks at 447 °C for the Pt/In₂O₃ catalyst and 427 °C for the pure In₂O₃ catalyst belong to CO₂ adsorption in the thermally induced oxygen vacancies.^{14,15,18} These results confirm that the total amount of the surface oxygen vacancies on the Pt/In₂O₃ catalyst is smaller than that of the pure In₂O₃ catalyst after hydrogen reduction, which is consistent with the results of the Raman spectra above.

UV-vis absorption spectra

Fig. 7a shows the UV-vis absorption spectra of the Pt/In₂O₃ catalyst. Fresh In₂O₃ shows a strong UV absorption because of its semiconducting nature. The band gap can be determined from absorption spectra using the Tauc relationship. Based on the plot of (αhν)² versus the photoenergy, hν, in Fig. 7b, the band gap of the Pt/In₂O₃-BT sample is determined to be 2.8 eV. The band gap is reduced to 2.57 eV for the Pt/In₂O₃ catalyst. This result indicates that the creation of the oxygen

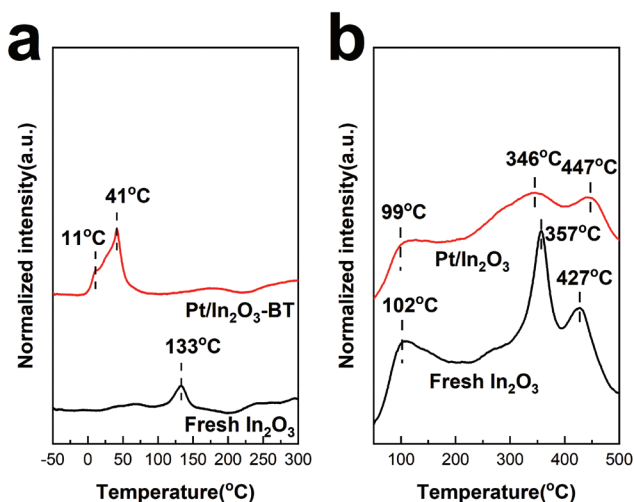


Fig. 6 (a) H₂-TPR profiles and (b) CO₂-TPD profiles of the Pt/In₂O₃ and pure In₂O₃ catalysts.

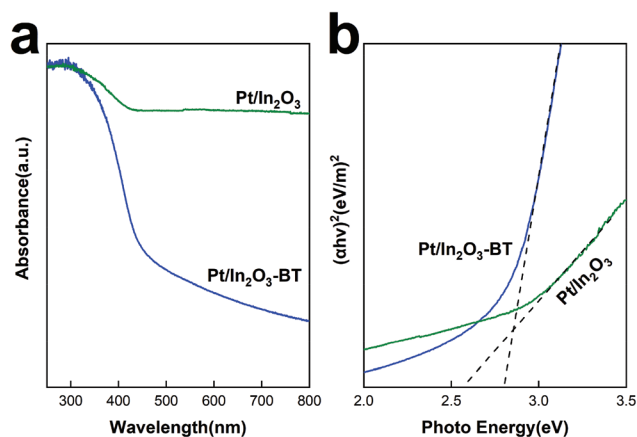


Fig. 7 (a) UV-vis absorption spectra and (b) Tauc plot of the Pt/In₂O₃ catalyst.

vacancies shifts the onset of light absorption from the UV to the visible region. It is also consistent with the Raman analyses discussed above: oxygen vacancies can be created over the Pt/In₂O₃ catalyst after hydrogen reduction. According to the literature,^{50,51} the increase in visible light absorption can be attributed to the electronic transition from the In₂O₃ valence band to the newly created oxygen vacancy level. The sensitivity of visible light absorption to the oxygen vacancies in In₂O₃ also makes Pt/In₂O₃ potentially useful in exploiting visible light for photocatalysts.

Conclusions

In this work, we prepared a novel Pt/In₂O₃ catalyst using the deposition-precipitation method. The catalyst was tested for methanol synthesis from CO₂ and hydrogen and the methanol selectivities of ca. 100% at temperatures below 225 °C, 74% at 275 °C and 54% at 300 °C on the Pt/In₂O₃ catalyst have been achieved. These results represent a significant improvement from the pure In₂O₃ catalyst and are the highest methanol selectivity reported on Pt catalysts. For example, the Pt/In₂O₃ catalyst shows a CO₂ conversion above 17% with a methanol formation rate of 0.542 g_{methanol} h⁻¹ g_{cat}⁻¹ at 300 °C and 5 MPa, while the CO₂ conversion is only 9.4% and the methanol formation rate is 0.335 g_{methanol} h⁻¹ g_{cat}⁻¹ for In₂O₃. In addition to the improved activity, the stability of the Pt/In₂O₃ catalyst is also outstanding. The methanol formation rate on the Pt/In₂O₃ catalyst after 5 h in the reaction stream only decreases to 95% of the initial rate and remains constant thereafter. This is significantly higher than 80% of the initial rate on the pure In₂O₃ catalyst reached after 9 h in the reaction stream. The textural characterization (by BET, XRD and TEM) of the catalysts indicates that the Pt nanoparticles are well-dispersed with an average particle size of 1.47 nm. More importantly, TEM and H₂-TPR showed that there were SMSIs between Pt and In₂O₃, resulting in the improved surface stability. In addition, the results of Raman spectra confirm that the

SMSI between Pt and In₂O₃ inhibits the over-reduction of In₂O₃. Consequently, the Pt/In₂O₃ catalyst showed a much higher catalytic activity and better stability than the pure In₂O₃ catalyst.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the National Key Research and Development Program of China (2016YFB0600902).

Notes and references

- 1 M.-Y. Wang, N. Wang, X.-F. Liu, C. Qiao and L.-N. He, *Green Chem.*, 2018, **20**, 1564–1570.
- 2 M. Lu, J. Zhang, Y. Yao, J. Sun, Y. Wang and H. Lin, *Green Chem.*, 2018, **20**, 4292–4298.
- 3 Z. Zhang, X.-Y. Zhou, J.-G. Wu, L. Song and D.-G. Yu, *Green Chem.*, 2020, **22**, 28–32.
- 4 J. Zhang, Q. Qian, Y. Wang, B. B. Asare Bediako, M. Cui, G. Yang and B. Han, *Green Chem.*, 2019, **21**, 233–237.
- 5 J. Albo, M. Alvarez-Guerra, P. Castaño and A. Irabien, *Green Chem.*, 2015, **17**, 2304–2324.
- 6 G. A. Olah, *Angew. Chem., Int. Ed.*, 2005, **44**, 2636–2639.
- 7 K. Räuchle, L. Plass, H.-J. Wernicke and M. Bertau, *Energy Technol.*, 2016, **4**, 193–200.
- 8 X. Dong, F. Li, N. Zhao, F. Xiao, J. Wang and Y. Tan, *Appl. Catal., B*, 2016, **191**, 8–17.
- 9 H. Ban, C. Li, K. Asami and K. Fujimoto, *Catal. Commun.*, 2014, **54**, 50–54.
- 10 S. Chen, J. Zhang, P. Wang, X. Wang, F. Song, Y. Bai, M. Zhang, Y. Wu, H. Xie and Y. Tan, *ChemCatChem*, 2019, **11**, 1448–1457.
- 11 J. Wang, G. Li, Z. Li, C. Tang, Z. Feng, H. An, H. Liu, T. Liu and C. Li, *Sci. Adv.*, 2017, **3**.
- 12 F. Studt, I. Sharafutdinov, F. Abild-Pedersen, C. F. Elkjær, J. S. Hummelshøj, S. Dahl, I. Chorkendorff and J. K. Nørskov, *Nat. Chem.*, 2014, **6**, 320.
- 13 F. Liao, X.-P. Wu, J. Zheng, M. M.-J. Li, A. Kroner, Z. Zeng, X. Hong, Y. Yuan, X.-Q. Gong and S. C. E. Tsang, *Green Chem.*, 2017, **19**, 270–280.
- 14 J. Ye, C. Liu and Q. Ge, *J. Phys. Chem. C*, 2012, **116**, 7817–7825.
- 15 J. Ye, C. Liu, D. Mei and Q. Ge, *ACS Catal.*, 2013, **3**, 1296–1306.
- 16 J. Ye, C.-j. Liu, D. Mei and Q. Ge, *J. Catal.*, 2014, **317**, 44–53.
- 17 K. Sun, Z. Fan, J. Ye, J. Yan, Q. Ge, Y. Li, W. He, W. Yang and C.-j. Liu, *J. CO₂ Util.*, 2015, **12**, 1–6.
- 18 N. Rui, Z. Wang, K. Sun, J. Ye, Q. Ge and C.-j. Liu, *Appl. Catal., B*, 2017, **218**, 488–497.

- 19 A. Tsoukalou, P. M. Abdala, D. Stoian, X. Huang, M. G. Willinger, A. Fedorov and C. R. Muller, *J. Am. Chem. Soc.*, 2019, **141**, 13497–13505.
- 20 M. S. Frei, M. Capdevila-Cortada, R. García-Muelas, C. Mondelli, N. López, J. A. Stewart, D. C. Ferré and J. Pérez-Ramírez, *J. Catal.*, 2018, **361**, 313–321.
- 21 M. Zhang, W. Wang and Y. Chen, *Appl. Surf. Sci.*, 2018, **434**, 1344–1352.
- 22 M. S. Frei, C. Mondelli, A. Cesarini, F. Krumeich, R. Hauert, J. A. Stewart, D. C. Ferré and J. Pérez-Ramírez, *ACS Catal.*, 2020, **10**, 1133–1145.
- 23 T.-y. Chen, C. Cao, T.-b. Chen, X. Ding, H. Huang, L. Shen, X. Cao, M. Zhu, J. Xu, J. Gao and Y.-F. Han, *ACS Catal.*, 2019, **9**, 8785–8797.
- 24 M. Zhang, M. Dou and Y. Yu, *Appl. Surf. Sci.*, 2018, **433**, 780–789.
- 25 M. Zhang, M. Dou and Y. Yu, *Phys. Chem. Chem. Phys.*, 2017, **19**, 28917–28927.
- 26 Z. Shi, Q. Tan, C. Tian, Y. Pan, X. Sun, J. Zhang and D. Wu, *J. Catal.*, 2019, **379**, 78–89.
- 27 Z. Shi, Q. Tan and D. Wu, *AIChE J.*, 2018, **65**, 1047–1058.
- 28 A. Bavykina, I. Yarulina, A. J. Al Abdulghani, L. Gevers, M. N. Hedhili, X. Miao, A. R. Galilea, A. Pustovarenko, A. Dikhtiarenko, A. Cadiau, A. Aguilar-Tapia, J.-L. Hazemann, S. M. Kozlov, S. Oud-Chikh, L. Cavallo and J. Gascon, *ACS Catal.*, 2019, **9**, 6910–6918.
- 29 M. S. Frei, C. Mondelli, R. García-Muelas, K. S. Kley, B. Puértolas, N. López, O. V. Safonova, J. A. Stewart, D. C. Ferré and J. Pérez-Ramírez, *Nat. Commun.*, 2019, **10**, 3377.
- 30 J. Wang, K. Sun, X. Jia and C.-j. Liu, *Catal. Today*, 2020, DOI: 10.1016/j.cattod.2020.05.020.
- 31 X. Jia, K. Sun, J. Wang, C. Shen and C.-j. Liu, *J. Energy Chem.*, 2020, **50**, 409–415.
- 32 Y.-L. Men, Y. Liu, Q. Wang, Z.-H. Luo, S. Shao, Y.-B. Li and Y.-X. Pan, *Chem. Eng. Sci.*, 2019, **200**, 167–175.
- 33 Y. Wang, J. Zhao, Y. Li and C. Wang, *Appl. Catal., B*, 2018, **226**, 544–553.
- 34 E. S. Gutterød, A. Lazzarini, T. Fjermestad, G. Kaur, M. Manzoli, S. Bordiga, S. Svelle, K. P. Lillerud, E. Skúlason, S. Øien-Ødegaard, A. Nova and U. Olsbye, *J. Am. Chem. Soc.*, 2020, **142**, 999–1009.
- 35 S. B. T. Tran, H. Choi, S. Oh and J. Y. Park, *Catal. Lett.*, 2019, **149**, 2823–2835.
- 36 B. Ouyang, S. Xiong, Y. Zhang, B. Liu and J. Li, *Appl. Catal., A*, 2017, **543**, 189–195.
- 37 Z. Wang, L. Huang, B. Su, J. Xu, Z. Ding and S. Wang, *Chem. – Eur. J.*, 2020, **26**, 517–523.
- 38 S. Kattel, B. Yan, J. G. Chen and P. Liu, *J. Catal.*, 2016, **343**, 115–126.
- 39 A. Sápi, G. Halasi, J. Kiss, D. G. Dobó, K. L. Juhász, V. J. Kolcsár, Z. Ferencz, G. Vári, V. Matolin, A. Erdőhelyi, Á. Kukovecz and Z. Kónya, *J. Phys. Chem. C*, 2018, **122**, 5553–5565.
- 40 K.-P. Yu, W.-Y. Yu, M.-C. Kuo, Y.-C. Liou and S.-H. Chien, *Appl. Catal., B*, 2008, **84**, 112–118.
- 41 E. S. Gutterød, S. Øien-Ødegaard, K. Bossers, A.-E. Nieuwelink, M. Manzoli, L. Braglia, A. Lazzarini, E. Borfecchia, S. Ahmadigoltapeh, B. Bouchevreau, B. T. Lønstad-Bleken, R. Henry, C. Lamberti, S. Bordiga, B. M. Weckhuysen, K. P. Lillerud and U. Olsbye, *Ind. Eng. Chem. Res.*, 2017, **56**, 13206–13218.
- 42 S. Dang, H. Yang, P. Gao, H. Wang, X. Li, W. Wei and Y. Sun, *Catal. Today*, 2019, **330**, 61–75.
- 43 T. Toyao, S. Kayamori, Z. Maeno, S. M. A. H. Siddiki and K.-i. Shimizu, *ACS Catal.*, 2019, **9**, 8187–8196.
- 44 T. Bielz, H. Lorenz, P. Amann, B. Klötzer and S. Penner, *J. Phys. Chem. C*, 2011, **115**, 6622–6628.
- 45 T. Bielz, H. Lorenz, W. Jochum, R. Kaindl, F. Klauser, B. Klötzer and S. Penner, *J. Phys. Chem. C*, 2010, **114**, 9022–9029.
- 46 D. R. Mullins and K. Z. Zhang, *Surf. Sci.*, 2002, **513**, 163–173.
- 47 Z. Yan, Z. Xu, J. Yu and M. Jaroniec, *Appl. Catal., B*, 2016, **199**, 458–465.
- 48 G. D. Gesesse, C. Wang, B. K. Chang, S.-H. Tai, P. Beaunier, R. Wojcieszak, H. Remita, C. Colbeau-Justin and M. N. Ghazzal, *Nanoscale*, 2020, **12**, 7011–7023.
- 49 H. Iida and A. Igarashi, *Appl. Catal., A*, 2006, **298**, 152–160.
- 50 J. Gan, X. Lu, J. Wu, S. Xie, T. Zhai, M. Yu, Z. Zhang, Y. Mao, S. C. I. Wang, Y. Shen and Y. Tong, *Sci., Rep.*, 2013, **3**, 1021.
- 51 F. Lei, Y. Sun, K. Liu, S. Gao, L. Liang, B. Pan and Y. Xie, *J. Am. Chem. Soc.*, 2014, **136**, 6826–6829.
- 52 H. Zhu, X. Wang, F. Yang and X. Yang, *Cryst. Growth Des.*, 2008, **8**, 950–956.
- 53 M. Kaur, N. Jain, K. Sharma, S. Bhattacharya, M. Roy, A. K. Tyagi, S. K. Gupta and J. V. Yakhmi, *Sens. Actuators, B*, 2008, **133**, 456–461.
- 54 O. Martin, A. J. Martín, C. Mondelli, S. Mitchell, T. F. Segawa, R. Hauert, C. Drouilly, D. Curulla-Ferré and J. Pérez-Ramírez, *Angew. Chem., Int. Ed.*, 2016, **55**, 6261–6265.
- 55 L. Wang, M. Ghoussoub, H. Wang, Y. Shao, W. Sun, A. A. Tountas, T. E. Wood, H. Li, J. Y. Y. Loh, Y. Dong, M. Xia, Y. Li, S. Wang, J. Jia, C. Qiu, C. Qian, N. P. Kherani, L. He, X. Zhang and G. A. Ozin, *Joule*, 2018, **2**, 1369–1381.
- 56 W. Wang, Z. Qu, L. Song and Q. Fu, *J. Energy Chem.*, 2020, **47**, 18–28.
- 57 S. M. Lee, H. Eom and S. S. Kim, *Environ. Technol.*, 2019, **1**–11.
- 58 Z. Yang, J. Li, H. Zhang, Y. Yang, M. Gong and Y. Chen, *Catal. Sci. Technol.*, 2015, **5**, 2358–2365.
- 59 R. Prins, *Chem. Rev.*, 2012, **112**, 2714–2738.
- 60 X. Jiang, X. Nie, Y. Gong, C. M. Moran, J. Wang, J. Zhu, H. Chang, X. Guo, K. S. Walton and C. Song, *J. Catal.*, 2020, **383**, 283–296.