

Cite this: *J. Mater. Chem. A*, 2019, 7, 20696

## Hydrogen evolution activity tuning *via* two-dimensional electron accumulation at buried interfaces†

Yudong Xue,<sup>id</sup>\*<sup>abcd</sup> Zachary S. Fishman,<sup>id</sup><sup>ac</sup> Yunting Wang,<sup>e</sup> Zhenhua Pan,<sup>ac</sup> Xin Shen,<sup>ac</sup> Rito Yanagi,<sup>ac</sup> Gregory S. Hutchings,<sup>a</sup> Mingzhao Liu,<sup>id</sup><sup>f</sup> Shili Zheng,<sup>b</sup> Yi Zhang,<sup>b</sup> Eric I. Altman,<sup>id</sup><sup>a</sup> and Shu Hu,<sup>id</sup><sup>\*ac</sup>

Developing efficient earth-abundant transition metal-based electrocatalysts for the hydrogen evolution reaction (HER) is crucial for hydrogen production at scale. This paper reports that the buried electrocatalytic interfaces between Ni–Fe sulfide (NiFeS) nanosheets and TiO<sub>2</sub> conformal coatings (about 5 nm) achieved remarkable HER activity improvement, lowering the HER overpotential from –170 mV to –107 mV at –50 mA cm<sup>–2</sup> in a base. Non-HER active, permeable TiO<sub>2</sub> coatings grown by atomic layer deposition (ALD) achieved continuous fine-tuning of the electronic properties at the buried TiO<sub>2</sub>/NiFeS interfaces, as a novel strategy and the main factor for electron accumulation at the interface. Core-level and valence band X-ray photoelectron spectroscopy (XPS) was used to investigate the TiO<sub>2</sub> electronic-structure tuning effect on the charge-transfer energetics during the HER. Their alkaline HER mechanism was elucidated by supplementing characterizations of membrane permeation, Tafel slope, and synchrotron X-ray absorption spectroscopy, which verified that the buried TiO<sub>2</sub>/NiFeS interfaces are electrocatalytically active. This study offers a general strategy for improving the charge-transfer kinetics of an electrocatalytic system by confining catalysis at a permeable solid–solid interface. The broad applicability of permeable and tunable coatings potentially accelerates the optimization of earth-abundant catalysts to achieve high performance under operationally relevant conditions.

Received 3rd July 2019  
Accepted 19th August 2019

DOI: 10.1039/c9ta07123g

rsc.li/materials-a

## Introduction

Hydrogen has been widely regarded as one of the most promising energy carriers to fulfill our need for a clean and sustainable future.<sup>1–4</sup> Broadly, low- and high-temperature advanced electrolysis, the solar thermochemical process, and photoelectrochemical (PEC) water splitting are considered three scalable hydrogen generation processes.<sup>5,6</sup> Both photoelectrochemical (PEC) water splitting and photovoltaic-driven electrolysis require electrocatalysts to produce hydrogen.<sup>7,8</sup> Among these technological pathways, development of efficient

and cost-effective electrocatalysts for the hydrogen evolution reaction (HER) is critical.<sup>9–16</sup> Moreover, while the HER in bases generally shows inferior activity than in acids because surface \*OH and \*H species compete for the active sites, numerous combinations of earth-abundant multi-element catalysts exist in bases, offering vast space for tunability.<sup>17</sup> To date, platinum-group metals and oxides, such as Pt and RuO<sub>2</sub>, have been among the most efficient, but their disadvantages are high cost and low abundance.<sup>18–20</sup> Alternatively, transition metal (Ni, Fe, Co, W, and Mo) compounds, such as metal sulfides, selenides, phosphides, nitrides, and carbides, have been proposed as efficient and low-cost alternatives to noble metals.<sup>21–26</sup>

Among them, transition metal sulfides have attracted special interest as HER catalysts due to their promising efficiency in alkaline media and their ease of synthesis.<sup>27,28</sup> Several early reports suggested that these drawbacks may be overcome by tuning surface atomic structures or confining molecular-scale intermediates; recent work has demonstrated that the HER activity of transition metal sulfides can be improved by introducing a synergistic constituent such as forming a bi-metallic sulfide.<sup>28–31</sup> Despite this development, the performance of transition metal sulfides is still far inferior to that needed for electrolysis and photolysis.

<sup>a</sup>Department of Chemical and Environmental Engineering, Yale University, New Haven, Connecticut 06511, USA. E-mail: ydxueipe@hotmail.com; shu.hu@yale.edu

<sup>b</sup>National Engineering Laboratory for Hydrometallurgical Cleaner Production Technology, CAS Key Laboratory of Green Process and Engineering, Institute of Process Engineering, Chinese Academy of Sciences, Beijing 100190, China

<sup>c</sup>Energy Sciences Institute, Yale University, West Haven, Connecticut 06516, USA

<sup>d</sup>University of Chinese Academy of Sciences, Beijing 100049, China

<sup>e</sup>School of Chemical and Environmental Engineering, China University of Mining and Technology (Beijing), Beijing 100083, China

<sup>f</sup>Center for Functional Nanomaterials, Brookhaven National Laboratory, Upton, New York 11973, USA

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ta07123g

A more recent report demonstrated tuning the interfacial electronic structures for improving electrocatalytic activity.<sup>32</sup> In many cases, tuning of catalysts does not need dramatic changes in their surface structures or compositions, such as in the case of bi-metallic alloying. Surface structure or composition tuning will change electronic structures, a conventional means for catalytic tuning.<sup>33</sup> Without explicitly tuning their structures or compositions, fine tuning of electronic structures for catalytic surfaces or interfaces is a new strategy. This paper is about further improvement of hydrogen evolution performance based on an optimized structure and composition of bimetallic catalysts.<sup>34</sup> This strategy offers an orthogonal control to structural or compositional tuning and uniquely allows for fine tuning on a continuous scale beyond conventional bi-metallic alloying. A small variation in spectroscopic signatures is expected, and will be correlated with observable changes in catalytic performance. Preliminary work showed that non-active TiO<sub>2</sub> coatings of up to 40 nm improved hydrogen evolution,<sup>29</sup> oxygen evolution, and chlorine evolution activities.<sup>35</sup> But their enhancement lacks clear elucidation: especially for the case of 40 nm coatings, there has been no convincing evidence to prove that there is enough reactant and product permittivity to keep up with the HER rate.

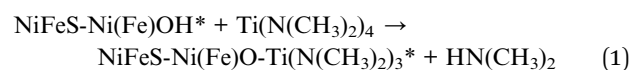
Atomic layer deposition (ALD) is a coating process that yields ultrathin overlayers with atomically precise thicknesses and exceptional conformality on high surface-area materials.<sup>34,36,37</sup> Several recent studies have shown that a porous coating can selectively block reactants.<sup>38</sup> Labrador *et al.* reported that a SiO<sub>x</sub>-coated Pt catalyst showed high selectivity for proton and H<sub>2</sub> transport against impurity electrodeposition and catalyst poisoning.<sup>39</sup> Vos *et al.* observed that MnO<sub>x</sub> thin films deposited on IrO<sub>2</sub> model catalysts increased the selectivity of oxygen evolution over chlorine evolution, acting as a diffusion barrier that prevents Cl<sup>-</sup> transport to active sites.<sup>40</sup> For PEC processes, impermeable, pinhole-free ALD TiO<sub>2</sub> coatings were used to prevent photo-corrosion in the pH range between 0 and 14, while being sufficiently transparent to reach the high quantum efficiency of protected semiconductors.<sup>37,41</sup> With a high throughput spatial technique, the ALD process is scalable for surface modification.<sup>42</sup> While coatings were normally grown thick enough to be impermeable, a less investigated technique is using ALD ultrathin membranes to enhance the electrocatalytic activity that bulk composition tuning cannot otherwise achieve. ALD coating is also considered favorable for stabilizing catalytic surfaces and preventing intermediates from leaving the lattice under operational conditions.<sup>43</sup>

In this study, we coated conformal, non-HER active, and permeable TiO<sub>2</sub> of 1–10 nm onto NiFe-bimetallic sulfide (NiFeS) nanosheet electrocatalysts by ALD. We exploited the effect of TiO<sub>2</sub>-coating/NiFeS-nanosheet buried hetero-junctions on their HER activity, and proposed a generic strategy for further boosting the HER activity of bi-metallic electrocatalysts. The structural and physicochemical properties, including crystalline structures, surface morphology, valence electronic structures, and interfacial band energetics, were systematically investigated. We discovered that the buried interface between TiO<sub>2</sub> and bimetallic sulfide nanosheets can be correlated with the favorably tuned electronic

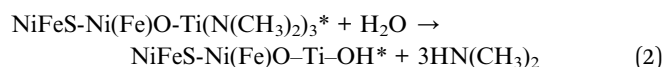
structures, suggesting a host of new opportunities for enhancing H<sub>2</sub> production at scale.

## Results and discussion

The TiO<sub>2</sub>/NiFeS hetero-junction catalyst was prepared *via* a “bottom-up” approach consisting of a hydrothermal process and sulfuration, followed by an ALD coating process (see the ESI† for details). During ALD, TiO<sub>2</sub> was chemically grown on the surface of NiFeS nanosheet/Ni foam using tetrakisdimethylamido-titanium (TDMAT) as the Ti precursor and H<sub>2</sub>O as the co-reactant. Ni foam was chosen as the catalyst support due to its three-dimensional internal pores with high alkali stability.<sup>44</sup> The pulse time of H<sub>2</sub>O and TDMAT was optimized to be 0.06 and 0.25 seconds, respectively, longer than that under previously reported conditions,<sup>37</sup> to achieve full surface coverage on NiFeS-loaded Ni foam. Fig. S1 and S2† illustrate the ALD process, which involves successive self-limiting surface reactions between the H<sub>2</sub>O and TDMAT over the NiFeS nanosheet substrate, indicating that there is enough precursor supplied to coat high surface-area materials conformally. According to the planar growth rate of 0.047 nm per cycle, the coating thicknesses were nominally 1, 2, 5, 7, and 10 nm, corresponding to 22, 43, 106, 149, and 213 ALD cycles. The surface reactions during the TDMAT exposure are expressed as eqn (1). Here, \* refers to surface reactive sites.



TDMAT can react with the hydroxyl groups that are formed on the NiFeS surface by a H<sub>2</sub>O pulse, and make Ti–O bonds by transferring protons (or possibly *via* protons attached to H<sub>2</sub>O molecules) to TDMAT's ligand complex to produce HN(CH<sub>3</sub>)<sub>2</sub> as a product which then leaves the surface. The subsequent surface reactions during the H<sub>2</sub>O exposure are expressed as eqn (2) and (3).



During the initial cycles, H<sub>2</sub>O will react with a surface-bound Ti(N(CH<sub>3</sub>)<sub>2</sub>)<sub>3</sub>\* to yield a hydroxyl-terminated \*Ti–OH surface, or H<sub>2</sub>O could bridge the two surface-(N(CH<sub>3</sub>)<sub>2</sub>)<sub>3</sub> sites that are close by to form a \*Ti–O–Ti\* oxo bridge. The surface chemistry of NiFeS before ALD growth was carefully controlled to minimize the formation of interfacial NiO<sub>x</sub> or FeO<sub>x</sub>, and the thermodynamics of TiO<sub>2</sub> compared to NiFeS native oxides favors the formation of NiFeS/TiO<sub>2</sub> interfaces but not NiO<sub>x</sub>- or FeO<sub>x</sub>-rich interfaces. After the initial cycles, TiO<sub>2</sub> was deposited on a TiO<sub>2</sub> surface until its deposition was completed. Unreacted precursor molecules, along with volatile products, were removed in a continuous flow of inert Ar carrier gas.

X-ray diffraction (XRD) was employed for structural characterization of the as-prepared samples and those coated with

varying ALD  $\text{TiO}_2$  thicknesses (1–10 nm), as shown in Fig. 1. Prior to the ALD process, the NiFeS nanosheet generated XRD patterns with diffraction peaks at  $21.7^\circ$ ,  $31.1^\circ$ ,  $37.8^\circ$ ,  $49.7^\circ$ , and  $55.4^\circ$ , which were consistent with the (101), (110), (003), (113), and (300) crystal planes of hexagonal  $\text{Ni}_3\text{S}_2$  (PDF no. 00-044-1418).<sup>45,46</sup> No diffraction peaks ascribed to iron sulfides can be observed, which indicates that the iron atoms are either substitutional sites in the lattice of  $\text{Ni}_3\text{S}_2$  or in an amorphous state that does not participate in X-ray scattering. In either case, the effect of iron on varying the lattice parameters is negligible. The chemical formula was  $\text{NiFe}_{1.29}\text{S}_{1.53}$ ; calculations are documented in the ESI.† After the ALD process, no additional diffraction peaks were introduced into the NiFeS structure. Moreover, the intensity of most  $\text{Ni}_3\text{S}_2$  peaks was maintained because the surface-deposited thin amorphous  $\text{TiO}_2$  barely attenuated the intensity of scattered X-rays from the underlying NiFeS crystallites.

The HER performance of the NiFeS and  $\text{TiO}_2/\text{NiFeS}$  electrocatalysts was compared by linear sweep voltammetry (LSV) measurements in a 1 M KOH (aq) electrolyte. As shown in Fig. 2a, the overpotentials at a current density of  $-50 \text{ mA cm}^{-2}$  were  $-141$ ,  $-133$ ,  $-107$ ,  $-132$ , and  $-185 \text{ mV}$  for the 1, 2, 5, 7, and 10 nm ALD  $\text{TiO}_2$  modifications, respectively. The reproducibility is good based on testing at least 5 times. The overpotential measurement error was  $\pm 5 \text{ mV}$ . Too high current densities and slow mass transport can cause current fluctuations, shown as spikes in Fig. 2a. Among these samples, the 5 nm ALD  $\text{TiO}_2/\text{NiFeS}$  exhibited the best HER performance, *i.e.*, the lowest overpotentials. To achieve a catalytic current density of  $-50 \text{ mA cm}^{-2}$ , an overpotential of only  $-107 \text{ mV}$  was needed for the 5 nm  $\text{TiO}_2/\text{NiFeS}$ , which is much lower than that of the as-synthesized NiFeS ( $-170 \text{ mV}$ ), NiFe-layered double hydroxide (NiFe-LDH,  $-326 \text{ mV}$ ), and Ni foam ( $-410 \text{ mV}$ , Fig. S3a†). This suggests that the HER activity of the NiFeS electrode can be tuned by coating catalysts with non-catalytic ALD films grown at

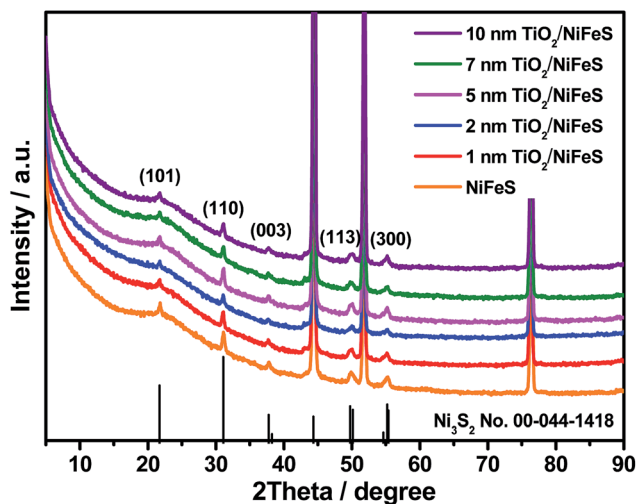


Fig. 1 X-ray diffraction (XRD) patterns of NiFeS nanosheets: bare sample and samples with 1, 2, 5, 7, and 10 nm of atomic layer deposition (ALD)  $\text{TiO}_2$  coating.

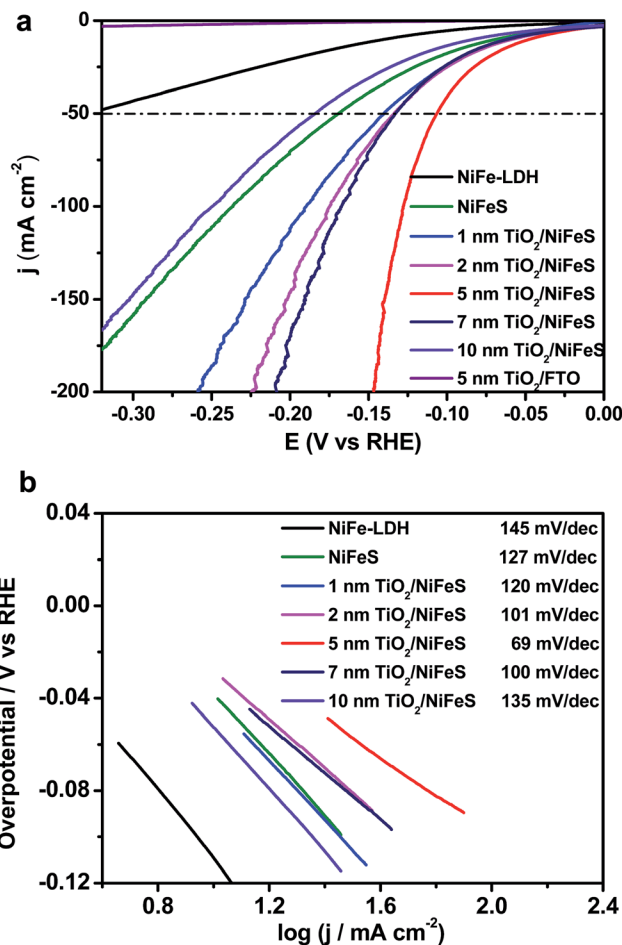


Fig. 2 Linear sweep voltammetry (LSV) curves (a) and the respective Tafel plots (b) of the ALD  $\text{TiO}_2/\text{NiFeS}$  electrodes and  $\text{TiO}_2/\text{FTO}$  (fluorine-doped tin oxide) electrodes measured in a 1 M KOH (aq) solution.

temperatures as low as  $150^\circ\text{C}$  without substantial variations in the original surface structures and compositions.  $\text{TiO}_2/\text{FTO}$  was chosen as the control to show that 5 nm  $\text{TiO}_2$  alone does not exhibit HER activity.

The electrocatalytic activity of the NiFeS electrode increases with increasing  $\text{TiO}_2$  layer thickness from 1 to 5 nm but decreases with increasing  $\text{TiO}_2$  layer thickness from 5 to 10 nm. This suggests that an optimal coating thickness is needed to achieve the most active NiFeS/ $\text{TiO}_2$ /electrolyte interface. The corresponding Tafel slope of the 5 nm  $\text{TiO}_2/\text{NiFeS}$  sample ( $69 \text{ mV dec}^{-1}$ ) was much smaller than that of the uncoated NiFeS ( $127 \text{ mV dec}^{-1}$ ), as shown in Fig. 2b. These results indicated that the 5 nm  $\text{TiO}_2$ -coated sample favored faster kinetics at a reduced Tafel slope, suggesting a possible shift in the rate-determining step; we will discuss this in detail in the catalytic mechanism section.

The 5 nm  $\text{TiO}_2/\text{NiFeS}$  interface was further characterized for elucidating the boosted HER activity. The overpotential of 5 nm  $\text{TiO}_2/\text{NiFeS}$  at  $50 \text{ mA cm}^{-2}$  was comparable to that of a reported Pt/C HER catalyst.<sup>47</sup> The turnover frequencies (TOFs)<sup>48</sup> for 5 nm  $\text{TiO}_2/\text{NiFeS}$  and uncoated NiFeS at a  $-100 \text{ mV}$  overpotential were 2.72 and  $0.56 \text{ s}^{-1}$ , respectively. The TOFs, with their

calculations elaborated in the ESI and shown in Table S1,<sup>†</sup> indicated that the intrinsic HER activity of TiO<sub>2</sub>/NiFeS is about five times better than that of the uncoated NiFeS catalyst (buried TiO<sub>2</sub>/NiFeS will later be proven to be active), which is comparable with that of the state-of-the-art, earth-abundant transition metal HER catalysts. As shown in Fig. S4,<sup>†</sup> electrochemical impedance spectroscopy (EIS) analysis indicated that after 5 nm amorphous TiO<sub>2</sub> was deposited onto the sulfide surface, the charge-transfer resistance of the TiO<sub>2</sub>/NiFeS electrode decreased slightly from 1.563 to 1.526 Ω. Moreover, the quantity of hydrogen accumulated at 10 mA cm<sup>-2</sup> based on the TiO<sub>2</sub>/NiFeS nanocomposites matched with the calculated amount for the HER assuming a faradaic efficiency of ~100% as shown in Fig. S5.<sup>†</sup> Their long-term HER performance testing at a fixed current density of 10 mA cm<sup>-2</sup> for 30 hours (Fig. S6<sup>†</sup>) showed excellent stability of overpotentials, better than that of the uncoated NiFeS for 10 hours.<sup>45,49</sup> Only the overpotential increased by 6 mV after 30 hours of operation. Besides, after HER testing, there was no change from the XRD pattern, as shown in Fig. S7.<sup>†</sup> The SEM image and XPS survey spectrum after the HER test, as shown in Fig. S8 and S9a,<sup>†</sup> indicated that TiO<sub>2</sub> was stable in alkaline media. Trace ion analysis by inductively coupled plasma-mass spectrometry (ICP-MS) further confirmed that, compared with NiFe-LDH and NiFeS, less Ni and Fe were dissolved for the TiO<sub>2</sub>/NiFeS samples during the HER (Table S2<sup>†</sup>).<sup>50</sup> These results also showed increased structural stability of the TiO<sub>2</sub>/NiFeS catalyst.

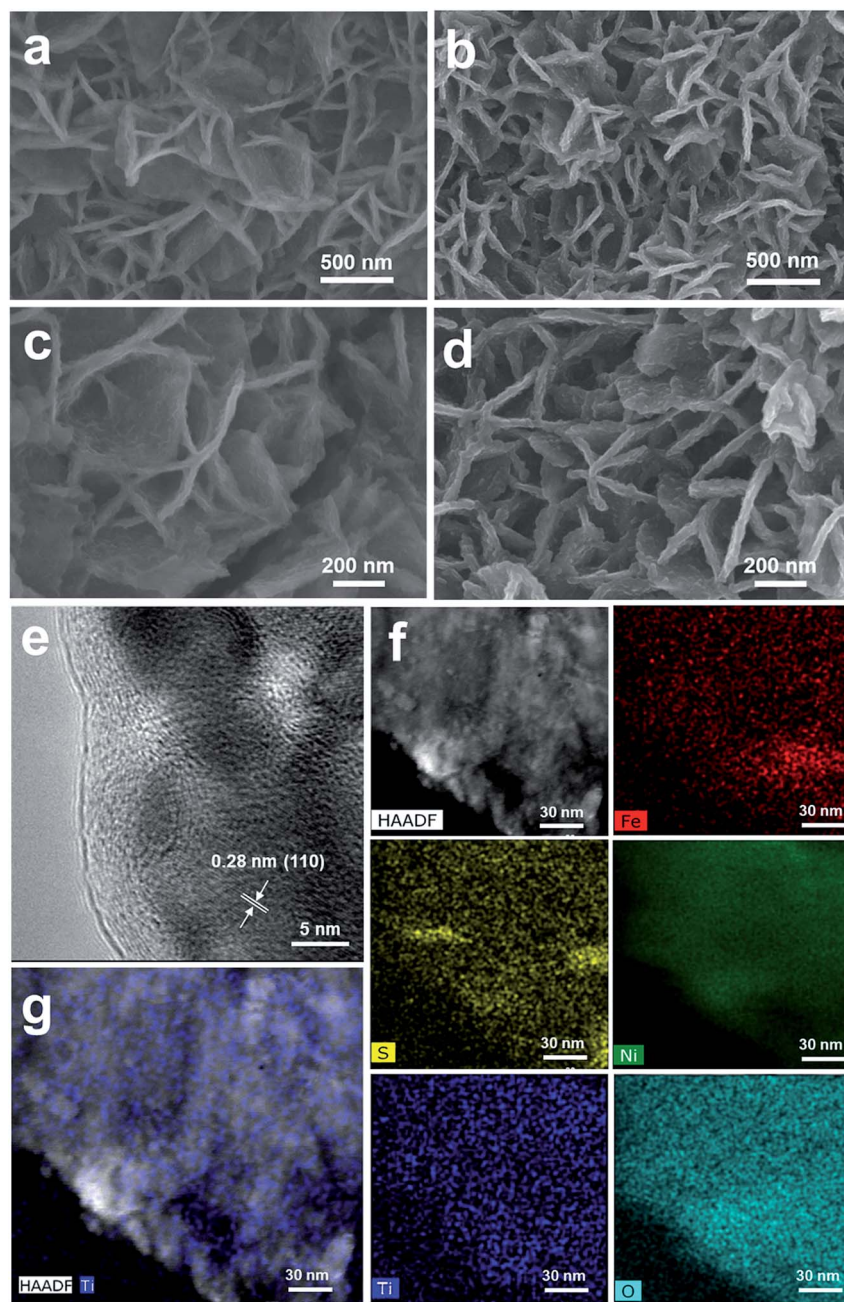
Fig. 3 shows the morphology and composition characterization of the as-synthesized NiFeS and 5 nm TiO<sub>2</sub>/NiFeS samples. In comparison to Fig. S10,<sup>†</sup> the nanosheet morphology of the NiFeS sample was well preserved after the sulfuration of NiFe-LDH. The original nanosheet structure morphology of TiO<sub>2</sub>-coated NiFeS (Fig. 3b and d) was well retained (Fig. 3a and c). The majority of the NiFeS surface was conformally coated by TiO<sub>2</sub> *via* the ALD process, *i.e.*, ~100% TiO<sub>2</sub> coverage. Moreover, comparing Fig. 3c and d, thicker nanosheets were observed for the 5 nm TiO<sub>2</sub>/NiFeS sample, indicating that the nanosheet surface was uniformly covered with amorphous TiO<sub>2</sub>. The full coverage of Ti and O elements over the NiFeS nanosheet surfaces was confirmed by scanning electron microscopy-energy dispersive X-ray spectroscopy (SEM-EDX) mapping (Fig. S11<sup>†</sup>). Furthermore, as shown in Fig. 3e, the high-resolution transmission electron microscopy (HRTEM) image exhibited lattice fringes with an interplanar spacing of 0.29 nm corresponding to the spacing of the (110) crystal planes of the NiFeS nanosheets. The HRTEM structural characterization indicated that the NiFeS nanosheet basal planes had a {001}-type crystal orientation, which was reported to be catalytically active.<sup>45</sup> Both the structural and elemental characterization experiments verified that an amorphous TiO<sub>2</sub> layer formed a heterojunction with the underlying NiFeS nanosheets with ~100% surface coverage on their basal planes (Fig. 3e). After ALD, the NiFeS crystal structure was maintained. As illustrated in the EDX elemental mapping images (Fig. 3f and g, and S12<sup>†</sup>), the Ni, Fe, and S elements were uniformly distributed within the NiFeS nanosheets, with the elemental intensity

varying with surface morphology occasionally, while the Ti and O elements covered the NiFeS surfaces uniformly.

The electronic structure of the buried TiO<sub>2</sub>/NiFeS heterojunction interfaces was characterized. X-ray photoelectron spectroscopy (XPS), including core-level (CL) and valence band (VB) spectra, was applied to construct the energy-band diagrams of the TiO<sub>2</sub>/NiFeS heterojunction interfaces.<sup>51</sup> All samples were ground from their surfaces, and TiO<sub>2</sub> was sufficiently thin for the X-ray to probe the buried TiO<sub>2</sub>/NiFeS interface. The measurement error for binding energies is below 0.03 eV. The XPS spectra of Ti 2p CL photoemission were deconvoluted as shown in Fig. 4a. In the TiO<sub>2</sub>/NiFeS sample, a small negative shift (0.2 eV) in the position of the Ti<sup>4+</sup> 2p<sub>3/2</sub> signal (from 458.55 to 458.35 eV) was observed, compared to TiO<sub>2</sub> on the fluorine-doped tin oxide (FTO) substrate (Fig. 4a). Valence band (VB) XPS data showed that TiO<sub>2</sub> stayed at the flat band after its deposition over FTO (by comparing the TiO<sub>2</sub> valence XPS spectrum on FTO with that on TiO<sub>2</sub> deposited on the NiFeS catalyst as shown in Fig. S9b<sup>†</sup>).<sup>37,51</sup> Because of the comparable valence band edges for TiO<sub>2</sub> on FTO and on NiFeS, the core-level peaks can be used to derive the band bending diagram at the NiFeS/TiO<sub>2</sub> interface (see the ESI<sup>†</sup> for detailed explanations). The negative shift indicated that the TiO<sub>2</sub> donated electrons to the NiFeS substrate across the buried interface. These observations indicated that NiFeS accepted electrons at the NiFeS/TiO<sub>2</sub> interface and the very interface became a thin, two-dimensional sheet of accumulated electrons. As shown in Fig. 4b, the TiO<sub>2</sub> O 1s CL photoemission peak-intensity shift was also negative 0.2 eV (from 529.90 to 529.70 eV), verifying the observed electron accumulation *via* interfacial charge transfer.<sup>52,53</sup> Accordingly, the increase in electron density at the NiFeS nanosheet surface is evident by a positive shift in the binding energies for Ni 2p, Fe 2p, and S 2p XPS peaks, as shown in Fig. 4c, d, and e, respectively. The chemical states of Ni<sup>2+</sup>, Fe<sup>2+</sup>, and S<sup>2-</sup> were maintained after coating. The direct evidence of electron accumulation is that the band edges and core-level peaks of NiFeS further shift downwards at heterojunction interfaces. Because of the TiO<sub>2</sub> coating, the electrons were accumulated at the buried interface, thus causing the band edges of NiFeS at the buried interface to be shifted upward by 0.5 eV relative to the band edges of the NiFeS bulk (for detailed calculations, see the ESI<sup>†</sup>).<sup>30,53</sup> VB XPS spectra of the TiO<sub>2</sub>/NiFeS and NiFeS samples are shown in Fig. S9c<sup>†</sup> and d,<sup>†</sup> respectively. The VB XPS data of the TiO<sub>2</sub>/NiFeS (Fig. S9c<sup>†</sup>) are comparable to the reported VBM value of "leaky" TiO<sub>2</sub> with a Ti<sup>3+</sup>-defect band inside its band gap.<sup>51</sup> The Fermi level of both samples is located at 0 eV. The valence-band maximum (VBM) position for the TiO<sub>2</sub> grown on NiFeS surfaces was assigned as 2.95 eV below the Fermi level of the heterojunction structure. The VB cut-off value of the NiFeS sample (Fig. S9d<sup>†</sup>) was at 0 eV, indicating that the NiFeS is either metallic or semi-metallic, which agrees with previous reports.<sup>45,46,54-56</sup>

Collectively, measurements of CL and VB X-ray photoemission spectra yielded a detailed diagram for interfacial band energetics, which is shown in Fig. 5. Despite the fact that the work function of sulfides is typically considered to be lower than that of TiO<sub>2</sub>, our measurements showed that the TiO<sub>2</sub>/NiFeS





**Fig. 3** Structural and elemental characterizations of the as-synthesized NiFeS (a and c) and 5 nm ALD TiO<sub>2</sub> grown on the NiFeS nanosheets by scanning electron microscopy (SEM) (b and d), and the high-resolution transmission electron microscopy (HRTEM) image (e), high-angle annual dark-field scanning transmission electron microscopy (HAADF-STEM) image, and energy dispersive X-ray spectroscopy (EDX) elemental mapping images of the 5 nm TiO<sub>2</sub>/NiFeS structure (f and g).

interfacial chemistry by ALD resulted in band edge alignment as shown in Fig. 5a. The density of states for NiFeS charge-transport bands is not as high as that of typical metals, because the CL peak positions did not shift with respect to the Fermi energy level. Therefore, the NiFeS material is considered a semi-metal rather than a conventional metal with much higher density of states.<sup>57,58</sup> In this case, electron accumulation lowered the band edge positions of NiFeS, relative to the Fermi level of TiO<sub>2</sub>/NiFeS heterojunction structures. The semi-metallic

properties of NiFeS ensure its downward band bending by 0.5 eV due to electron accumulation, while ALD TiO<sub>2</sub> yielded an upward band bending of 0.2 eV due to electron transfer from TiO<sub>2</sub> to NiFeS and oxygen vacancy-induced space charge. Therefore, the electrons at the buried interface were found to be delocalized and accumulated into a thin sheet and were distributed across TiO<sub>2</sub>-coated NiFeS surfaces, thus boosting reductive charge transfer. The electrons were estimated to accumulate within less than the 5 nm skin depth of NiFeS

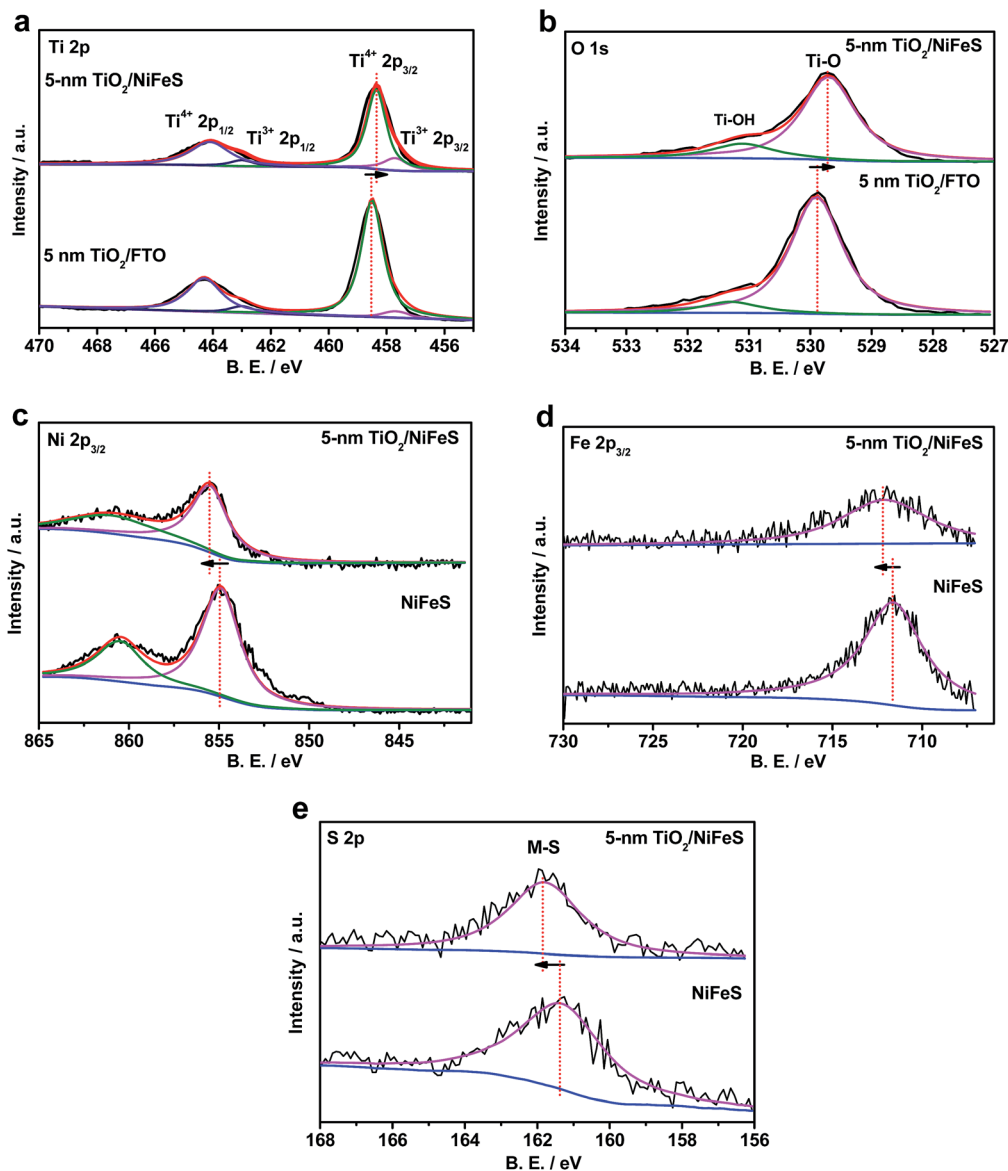


Fig. 4 X-ray photoelectron spectroscopy (XPS) spectra of Ti 2p (a) and O 1s (b) core-level photoemissions for the 5 nm ALD TiO<sub>2</sub> grown on the NiFeS and FTO substrates, and XPS spectra of Ni 2p (c), Fe 2p (d), and S 2p (e) core-level photoemissions comparing the 5 nm TiO<sub>2</sub>/NiFeS and bare NiFeS samples (blue represents the background. Black represents the raw data. The other colours represent fitting curves. The XPS measurement error for binding energies is below 0.03 eV).

surfaces, according to the band-edge shift and density of states. The relative position of NiFeS band edges that participate in charge transfer during the HER is raised, with respect to the potential of surface \*OH and \*H states. This tuning is not possible without the formation of a NiFeS/TiO<sub>2</sub> heterojunction interface. The band diagram analysis indicates that the present strategy is appropriate for tuning the catalytic properties of semi-metals and semiconductors.

The ALD TiO<sub>2</sub> also functions as a sponge that efficiently transfers OH<sup>-</sup>, water or protons to the buried interface.<sup>51,59</sup> It has been reported that metal oxides (especially TiO<sub>2</sub>) in the solution can transfer/adsorb the protons in the interfacial proton-coupled electron transfer reactions.<sup>59,60</sup> A previous study indicated that only H-adsorbed active sites, instead of the OH

bonds formed in the direct Volmer step, actively participate in the alkaline HER.<sup>61</sup> Therefore, it is important to recognize the proton permeation properties of ALD TiO<sub>2</sub> membranes in addition to their interfacial tuning effect. We constructed a solution-based membrane permeation test system (see Fig. S13<sup>†</sup> for the setup and Fig. S14<sup>†</sup> for the results). The original Si<sub>3</sub>N<sub>4</sub> film covered on the window of the TEM grid blocks the OH<sup>-</sup> transfer, which ensures that the pH value of the permeation-side remains below pH 7. With a hole and without the TiO<sub>2</sub> membrane, the concentration of OH<sup>-</sup> should increase to 0.005 mol L<sup>-1</sup> in the original permeate compartment of deionized (DI) water by mixing with the pH 12 solution in the original compartment of the OH<sup>-</sup> source. The pH measurement without the TiO<sub>2</sub> membrane is consistent with the above

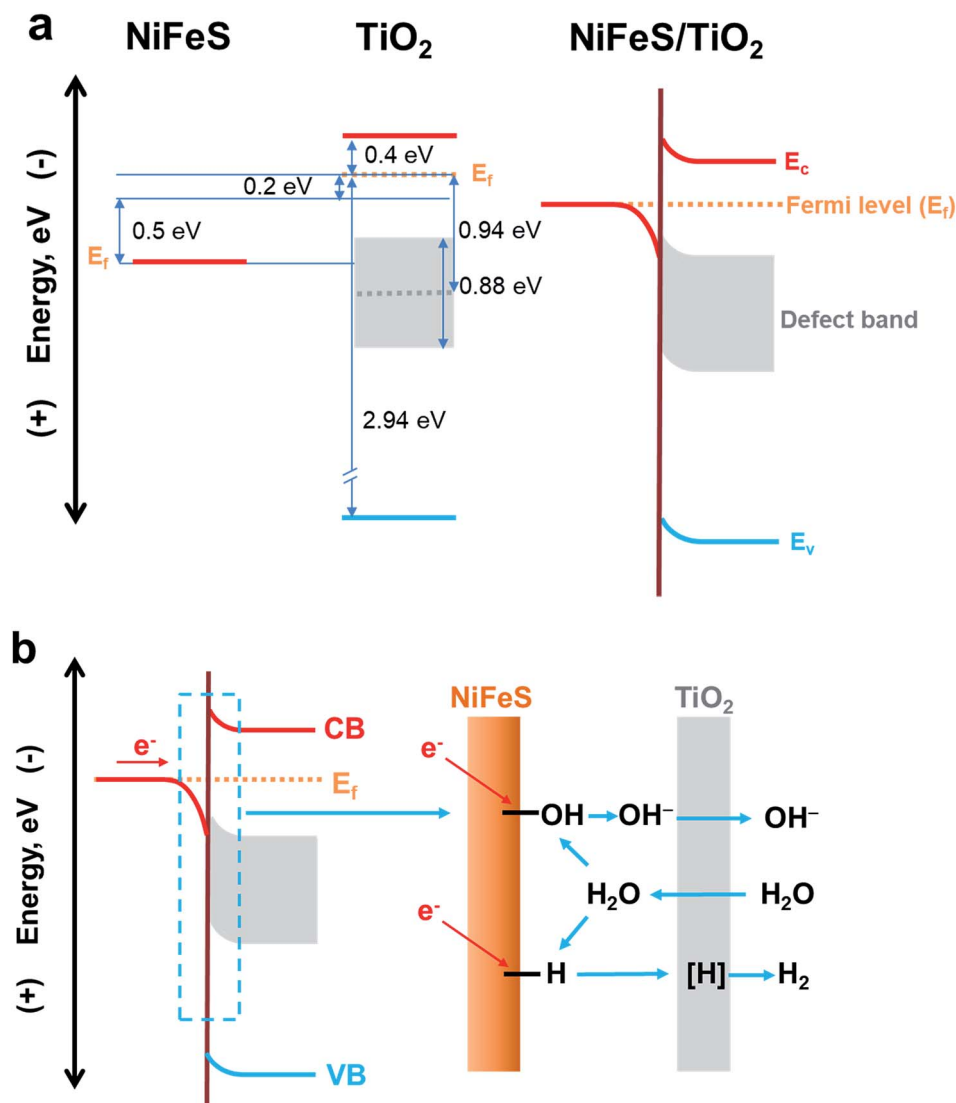


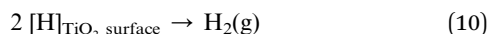
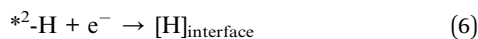
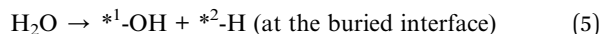
Fig. 5 (a) Energy-band diagrams of the TiO<sub>2</sub>/NiFeS interfaces. (b) Schematic illustration of the NiFeS/TiO<sub>2</sub> buried interface for boosting HER activity.

calculation, as the pH reached 11.66 after equilibration. With 5 nm TiO<sub>2</sub> covering the window (Si<sub>3</sub>N<sub>4</sub> dissolved in hot KOH leaving TiO<sub>2</sub> membranes only), the pH in the permeation-side increases to 10.95 ([OH<sup>-</sup>] = 0.0009 mol L<sup>-1</sup>), indicating that the 5 nm TiO<sub>2</sub> membrane presents considerable OH<sup>-</sup> permeation properties. After 18% OH<sup>-</sup> transfer from the original alkaline side of the membrane to the permeation side, excess amounts of K<sup>+</sup> and OH<sup>-</sup> will accumulate on the retentate and permeate sides of the membrane, respectively. The electrochemical potential will drop across the TiO<sub>2</sub> membrane, and the electric field formed will inhibit the OH<sup>-</sup> from further passing through the membrane. At steady-state, the permeation-side pH cannot reach 11.66 because the permittivity of K<sup>+</sup> is less than that of OH<sup>-</sup>. Our permeation experiment also showed that TiO<sub>2</sub> membranes can conduct protons, *i.e.*, H<sup>+</sup> (Fig. S14b†). It is reasonable to accept that atomic hydrogen intermediates, denoted as [H], which participated in the HER, can also be

transported through the as-fabricated TiO<sub>2</sub> semi-permeable membrane on the NiFeS surface. It is possible that the water molecules and OH<sup>-</sup> ions transported in and out of the membranes participate in the TiO<sub>2</sub>/NiFeS alkaline HER, because their sizes are in between those of H<sup>+</sup> and SO<sub>4</sub><sup>2-</sup>. The through-membrane transport properties for H<sup>+</sup>, OH<sup>-</sup>, H<sub>2</sub>O, and the electrolyte species and the associated catalytic mechanism are a subject for further study.

The electrochemical kinetics of the TiO<sub>2</sub>/NiFeS heterostructure were compared with those of the NiFe-LDH and NiFeS nanosheets. The Tafel slope (Fig. 2b) conveys important information about the rate-determining step (RDS) of the multi-step HER process. The decrease in Tafel slope as the TiO<sub>2</sub> coating thickness increases, in conjunction with TiO<sub>2</sub>'s permittivity to a host of catalytic species, supports the activity and tuning hypothesis for the buried catalytic interfaces. Without a buried interface, the HER in alkaline media involves three main steps:

the Volmer step ( $\text{H}_2\text{O} + \text{e}^- \rightarrow * \text{H} + \text{OH}^-$ ), the Heyrovsky step ( $* \text{H} + \text{H}_2\text{O} + \text{e}^- \rightarrow \text{H}_2 + \text{OH}^-$ ), and the Tafel step ( $* \text{H} + * \text{H} \rightarrow \text{H}_2$ ), where \* indicates surface active sites.<sup>45</sup> The corresponding Tafel slope of the Volmer, Heyrovsky, and Tafel steps as the RDS should be 120, 40, and 30  $\text{mV dec}^{-1}$ , respectively.<sup>17</sup> There is an obvious shift in the mechanism with and without the  $\text{TiO}_2$  modified overlayers: the Tafel slope decreases from 127 to 69  $\text{mV dec}^{-1}$ . When the coating thickness increases from 5 to 10 nm, the RDS becomes limited by  $[\text{H}]$  transport as the Tafel slope increases from 69 to 135  $\text{mV dec}^{-1}$ .<sup>62</sup> The mechanistic shift may be evidenced by the observed decrease in the Tafel slope from 120 to 69  $\text{mV dec}^{-1}$  with increasing  $\text{TiO}_2$  coating thickness (from 1 to 5 nm). Although the Tafel slopes for bare NiFeS and NiFeS with an optimal coating thickness appear similar, the mechanism behind these measured Tafel slopes with and without  $\text{TiO}_2$  overlayers is completely different. Therefore, we have proposed one plausible mechanism though we have not yet detected the postulated molecular species at the buried catalytic interface.



$\text{H}_2\text{O}$  dissociation can produce surface bonded  $*\text{-OH}$  and  $*\text{-H}$  at the buried interface, followed by a competitive electron-transfer process to  $*\text{-OH}$  vs. to  $*\text{-H}$  at the  $\text{Ti}^{3+}$ -coordinated, confined NiFeS catalytic surfaces. Then, atomic  $[\text{H}]$  will be generated after the reduction of  $*\text{-H}$  (step (6)). The  $[\text{H}]$  is considered to migrate from the interface to the catalyst/liquid interface through the 5 nm  $\text{TiO}_2$  coating. There is a correlation between the original NiFeS active sites buried by  $\text{TiO}_2$  coatings and their enhanced HER activity. It is postulated that the buried  $\text{TiO}_2/\text{NiFeS}$  interfaces are the main contributor for improving HER performance. Direct observation of catalytic intermediates and active sites at buried interfaces is an ongoing challenge for *in operando* characterization of electrochemical processes and may be done in the future.

The  $\text{TiO}_2/\text{NiFeS}$  buried interfaces also provide  $\text{Ti-OH}$  bonding sites, which may catch surface  $\text{OH}$  from NiFeS to further alleviate  $*\text{-OH}$  poisoning of the active sites and favor the Heyrovsky step. Additionally, the  $\text{TiO}_2$  after being deposited on NiFeS contains more low-coordinated Ti ions ( $\text{Ti}^{3+}$ ) than  $\text{TiO}_2$  on FTO substrates (satellite peaks next to  $\text{Ti}^{4+}$  peaks in Fig. 4a);<sup>63</sup> this condition is consistent with electronic filling into a  $\text{Ti}^{3+}$ -defect band.<sup>51</sup> It is reasonable to consider that the catalytically active sites extended into the  $\text{TiO}_2$  phase near the  $\text{TiO}_2/\text{NiFeS}$  buried interface, as  $\text{Ti}^{3+}$  ions were shown to present high HER

activity.<sup>63</sup> Therefore, the structural and energy-diagram characterization confirms that the tuning of NiFeS interfacial electrocatalytic properties *via* charge-transfer interaction and  $\text{Ti}^{3+}$  covalent coordination can be continuous and uniform.

At 10  $\text{mA cm}^{-2}$ , the stability after 30 hours of continuous operation verified the technological feasibility for photo-electrochemical water splitting. Promoting the out-diffusion of atomic hydrogen intermediates by increasing the  $\text{TiO}_2$  porosity may improve the electrochemical stability of NiFeS used in an electrolyzer. The coating is beneficial because it prevents the anticipated lattice exchange so that  $\text{Ni}^{2+}$  and  $\text{Fe}^{2+}$  cations are not liberated into the electrolyte. The Ni-H and Fe-H hydride intermediates may also favor leaving the NiFeS lattice with a relatively large equilibrium constant for their soluble species.<sup>43</sup> One of the multiple functions of  $\text{TiO}_2$  is the physical confinement of the lattice cations or surface hydrides to prevent dissolution. This aspect was confirmed from the ICP-MS results of dissolved  $\text{Ni}^{2+}$  and  $\text{Fe}^{2+}$  (Table S2†).  $\text{TiO}_2$  coating resulted in less dissolution of the active element into the electrolyte than in the case of uncoated NiFeS catalysts. So far, we have demonstrated that the rate of the alkaline HER was improved by facilitating the Heyrovsky step on the  $\text{TiO}_2$  membrane/coating, optimizing the  $[\text{H}]$  formation and transport at the  $\text{TiO}_2/\text{NiFeS}$  buried interface.

In order to better understand the atomic-scale structure of amorphous  $\text{TiO}_2$  membranes during electrocatalysis, the X-ray absorption spectra (XAS) of 5 nm  $\text{TiO}_2/\text{NiFeS}$  and  $\text{TiO}_2/\text{fused silica}$  ( $\text{TiO}_2/\text{silica}$ ) were acquired at the National Synchrotron Light Source II (NSLS II), shown in Fig. 6. Both  $\text{TiO}_2/\text{NiFeS}$  and  $\text{TiO}_2/\text{silica}$  gave only one peak in the pre-edge region near 4966 eV, confirming that the ALD  $\text{TiO}_2$  is amorphous on NiFeS and quartz.<sup>64</sup>  $\text{TiO}_2/\text{NiFeS}$  showed a slight decrease in the Ti absorption-edge energy, which is considered to be a consequence of the reduction of Ti cations in the  $\text{TiO}_2$  coating.<sup>65</sup>  $\text{TiO}_2$  on NiFeS catalysts showed a Ti coordination environment similar to that in the  $\text{TiO}_2/\text{silica}$ , as revealed in X-ray absorption near-edge structure (XANES) spectra. Two obvious XANES peaks

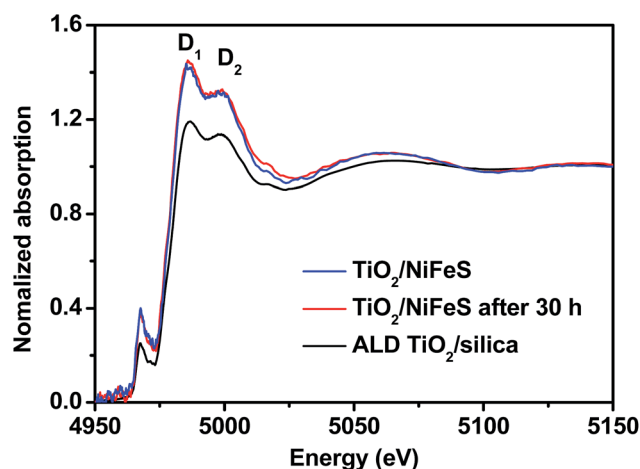


Fig. 6 Comparison of Ti K-edge X-ray absorption near edge structure (XANES) data collected on the as-prepared  $\text{TiO}_2/\text{silica}$  and 5 nm  $\text{TiO}_2/\text{NiFeS}$  nanosheets attached to Ni foam, before and after 30 hours of HER operation.



at 4985.8 eV and 4999.0 eV, marked  $D_1$  and  $D_2$ , were observed in both  $\text{TiO}_2/\text{silica}$  and  $\text{TiO}_2/\text{NiFeS}$ . In comparison, the ratio of  $D_1/D_2$  of the  $\text{TiO}_2/\text{NiFeS}$  (1.099) was significantly higher than that of  $\text{TiO}_2/\text{silica}$  (1.044). The larger  $D_1/D_2$  ratio indicates a stronger charge-transfer state contribution, which suggests more electron exchange between NiFeS and the  $\text{TiO}_2$  coating,<sup>66</sup> thereby resulting in the accumulation of electrons at the  $\text{TiO}_2/\text{NiFeS}$  interface. The XANES analysis indicated that the  $\text{TiO}_2$  layer was in a more reduced electronic state. XAS analysis is in accordance with the XPS results of the low valence ( $\text{Ti}^{3+}$ ) state in the coating. Moreover, comparing the  $\text{TiO}_2/\text{NiFeS}$  catalyst before and after the 30 hour HER operation, identical Ti K-edge XANES spectra were obtained, again indicating that the electronic structure of the  $\text{TiO}_2$  coating is stable after long-term operation.

## Conclusions

In summary,  $\text{TiO}_2/\text{NiFeS}$  heterostructured HER electrocatalysts were synthesized by atomic layer deposition (ALD) coating over hydrothermally grown NiFeS nanosheets. The  $\text{TiO}_2$  layers were uniformly coated over the surface of NiFeS, resulting in fine tuning of both the electronic properties of the buried  $\text{TiO}_2/\text{bimetallic sulfide}$  HER interfaces and the steric effects by  $\text{TiO}_2$ -coating confinement. The optimally designed 5 nm  $\text{TiO}_2/\text{NiFeS}$  heterostructures showed the highest HER activity, displaying a current density of  $-50 \text{ mA cm}^{-2}$  at a small overpotential of  $-107 \text{ mV}$ , a Tafel slope as small as  $69 \text{ mV dec}^{-1}$ , and 30 hour stability in alkaline media. We achieved small overpotentials and long-term stability for  $10 \text{ mA cm}^{-2}$ , which is an important benchmark for light-driven water splitting. This enhanced performance for bare NiFeS nanosheets can be explained by the favorable electronic structure at the buried interface: the charge-transfer driving force from NiFeS band edges to HER intermediates was increased. The ALD  $\text{TiO}_2$  coating is considered an electron donor and proton sponge. Observing a significant reduction of the Tafel slope from  $127 \text{ mV dec}^{-1}$  for bare NiFeS to  $69 \text{ mV dec}^{-1}$  for  $\text{TiO}_2/\text{NiFeS}$ , we proposed that the rate of the alkaline HER was improved by facilitating the Heyrovsky step at the buried interface, and optimizing the transport of hydrogen intermediates through the  $\text{TiO}_2$  membrane. The results of the X-ray absorption near-edge structure (XANES) further verified electronic interactions between NiFeS and the  $\text{TiO}_2$  coating. This study leads to a set of strategies including fine tuning the surface electronic properties of semi-metallic or semiconductive catalysts and confining charges and intermediates at a buried interface formed by a permeable coating. This approach has potential for general applicability in improving the performance of practical water splitting and other electrocatalytic reactions.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

We acknowledge the financial support of the University of Chinese Academy of Sciences (UCAS[2015]37) Joint PhD

Training Program, the National Natural Science Foundation of China (No. 51774261), and the Petroleum Research Fund (ND55524). This research used resources of the Center for Functional Nanomaterials, and beamline 8-ID ISS (Inner Shell Spectroscopy) at the National Synchrotron Light Source II; both are DOE Office of Science User Facilities operated for the DOE Office of Science by Brookhaven National Laboratory under Contract No. DE-SC00122704. The authors would like to thank Dr Min Li of Yale's Materials Characterization Core (MCC) and Dr Michael Rooks of the Yale Institute for Nanoscience and Quantum Engineering (YINQE) for their invaluable help.

## References

- 1 E. J. Popczun, C. G. Read, C. W. Roske, N. S. Lewis and R. E. Schaak, *Angew. Chem., Int. Ed.*, 2014, **53**, 5427–5430.
- 2 C. C. McCrory, S. Jung, I. M. Ferrer, S. M. Chatman, J. C. Peters and T. F. Jaramillo, *J. Am. Chem. Soc.*, 2015, **137**, 4347–4357.
- 3 E. J. Popczun, J. R. McKone, C. G. Read, A. J. Biacchi, A. M. Wiltrout, N. S. Lewis and R. E. Schaak, *J. Am. Chem. Soc.*, 2013, **135**, 9267–9270.
- 4 C. Xia, H. Liang, J. Zhu, U. Schwingenschlögl and H. N. Alshareef, *Adv. Energy Mater.*, 2017, **7**, 1602089.
- 5 I. Dincer and C. Acar, *Int. J. Hydrogen Energy*, 2015, **40**, 11094–11111.
- 6 M. Xiao, B. Luo, M. Lyu, S. Wang and L. Wang, *Adv. Energy Mater.*, 2018, **8**, 1701605.
- 7 J. Chen, X. J. Wu, L. Yin, B. Li, X. Hong, Z. Fan, B. Chen, C. Xue and H. Zhang, *Angew. Chem., Int. Ed.*, 2015, **54**, 1210–1214.
- 8 P. Zhang, L. Yu and X. W. D. Lou, *Angew. Chem., Int. Ed.*, 2018, **57**, 15076–15080.
- 9 G. Zhang, G. Wang, Y. Liu, H. Liu, J. Qu and J. Li, *J. Am. Chem. Soc.*, 2016, **138**, 14686–14693.
- 10 T. Zheng, W. Sang, Z. He, Q. Wei, B. Chen, H. Li, C. Cao, R. Huang, X. Yan, B. Pan, S. Zhou and J. Zeng, *Nano Lett.*, 2017, **17**, 7968–7973.
- 11 R. Miao, B. Dutta, S. Sahoo, J. He, W. Zhong, S. A. Cetegen, T. Jiang, S. P. Alpay and S. L. Suib, *J. Am. Chem. Soc.*, 2017, **139**, 13604–13607.
- 12 X. Wang, W. Ma, Z. Xu, H. Wang, W. Fan, X. Zong and C. Li, *Nano Energy*, 2018, **48**, 500–509.
- 13 Y. Qiu, Z. Wen, C. Jiang, X. Wu, R. Si, J. Bao, Q. Zhang, L. Gu, J. Tang and X. Guo, *Small*, 2019, **15**, 1900014.
- 14 X. Du, W. Lian and X. Zhang, *Int. J. Hydrogen Energy*, 2018, **43**, 20627–20635.
- 15 X. Du, Q. Wang, Y. Li and X. Zhang, *Dalton Trans.*, 2018, **47**, 10273–10280.
- 16 X. Du, X. Zhang, Y. Li and M. Zhao, *Int. J. Hydrogen Energy*, 2018, **43**, 19955–19964.
- 17 N. Mahmood, Y. Yao, J.-W. Zhang, L. Pan, X. Zhang and J.-J. Zou, *Adv. Sci.*, 2018, **5**, 1700464.
- 18 X. P. Yin, H. J. Wang, S. F. Tang, X. L. Lu, M. Shu, R. Si and T. B. Lu, *Angew. Chem., Int. Ed.*, 2018, **57**, 9382–9386.
- 19 J. N. Tiwari, S. Sultan, C. W. Myung, T. Yoon, N. Li, M. Ha, A. M. Harzandi, H. J. Park, D. Y. Kim,

- S. S. Chandrasekaran, W. G. Lee, V. Vij, H. Kang, T. J. Shin, H. S. Shin, G. Lee, Z. Lee and K. S. Kim, *Nat. Energy*, 2018, **3**, 773–782.
- 20 X. Du, Z. Yang, Y. Li, Y. Gong and M. Zhao, *J. Mater. Chem. A*, 2018, **6**, 6938–6946.
- 21 X. Li, W. Liu, M. Zhang, Y. Zhong, Z. Weng, Y. Mi, Y. Zhou, M. Li, J. J. Cha, Z. Tang, H. Jiang, X. Li and H. Wang, *Nano Lett.*, 2017, **17**, 2057–2063.
- 22 Z. W. Seh, K. D. Fredrickson, B. Anasori, J. Kibsgaard, A. L. Strickler, M. R. Lukatskaya, Y. Gogotsi, T. F. Jaramillo and A. Vojvodic, *ACS Energy Lett.*, 2016, **1**, 589–594.
- 23 L.-A. Stern, L. Feng, F. Song and X. Hu, *Energy Environ. Sci.*, 2015, **8**, 2347–2351.
- 24 Y. Zhou, J. V. Pondick, J. L. Silva, J. M. Woods, D. J. Hynek, G. Matthews, X. Shen, Q. Feng, W. Liu, Z. Lu, Z. Liang, B. Brena, Z. Cai, M. Wu, L. Jiao, S. Hu, H. Wang, C. M. Araujo and J. J. Cha, *Small*, 2019, **15**, e1900078.
- 25 H. Zhou, F. Yu, Y. Huang, J. Sun, Z. Zhu, R. J. Nielsen, R. He, J. Bao, W. A. Goddard III, S. Chen and Z. Ren, *Nat. Commun.*, 2016, **7**, 12765.
- 26 Y. Li, H. Zhang, M. Jiang, Q. Zhang, P. He and X. Sun, *Adv. Funct. Mater.*, 2017, **27**, 1702513.
- 27 X. Y. Yu and X. W. Lou, *Adv. Energy Mater.*, 2018, **8**, 1701592.
- 28 S.-Y. Huang, D. Sodano, T. Leonard, S. Luiso and P. S. Fedkiw, *J. Electrochem. Soc.*, 2017, **164**, F276–F282.
- 29 C. Bae, T. A. Ho, H. Kim, S. Lee, S. Lim, M. Kim, H. Yoo, J. M. Montero-Moreno, J. H. Park and H. Shin, *Sci. Adv.*, 2017, **3**, e1602215.
- 30 Y. Wu, X. Liu, D. Han, X. Song, L. Shi, Y. Song, S. Niu, Y. Xie, J. Cai, S. Wu, J. Kang, J. Zhou, Z. Chen, X. Zheng, X. Xiao and G. Wang, *Nat. Commun.*, 2018, **9**, 1425.
- 31 X. Long, G. Li, Z. Wang, H. Zhu, T. Zhang, S. Xiao, W. Guo and S. Yang, *J. Am. Chem. Soc.*, 2015, **137**, 11900–11903.
- 32 C. Chu, D. Huang, Q. Zhu, E. Stavitski, J. A. Spies, Z. Pan, J. Mao, H. L. Xin, C. A. Schmuttenmaer, S. Hu and J.-H. Kim, *ACS Catal.*, 2018, **9**, 626–631.
- 33 Y. Huang, J. Hu, H. Xu, W. Bian, J. Ge, D. Zang, D. Cheng, Y. Lv, C. Zhang, J. Gu and Y. Wei, *Adv. Energy Mater.*, 2018, **8**, 1800789.
- 34 Y. Xue, Z. S. Fishman, J. A. Röhr, Z. Pan, Y. Wang, C. Zhang, S. Zheng, Y. Zhang and S. Hu, *J. Mater. Chem. A*, 2018, **6**, 21918–21926.
- 35 C. E. Finke, S. T. Omelchenko, J. T. Jasper, M. F. Lichterman, C. G. Read, N. S. Lewis and M. R. Hoffmann, *Energy Environ. Sci.*, 2019, **12**, 358–365.
- 36 N. Cheng, Y. Shao, J. Liu and X. Sun, *Nano Energy*, 2016, **29**, 220–242.
- 37 S. Hu, M. R. Shaner, J. A. Beardslee, M. Lichterman, B. S. Brunshwig and N. S. Lewis, *Science*, 2014, **344**, 1005–1009.
- 38 D. V. Esposito, *ACS Catal.*, 2017, **8**, 457–465.
- 39 N. Y. Labrador, E. L. Songcuan, C. De Silva, H. Chen, S. J. Kurdziel, R. K. Ramachandran, C. Detavernier and D. V. Esposito, *ACS Catal.*, 2018, **8**, 1767–1778.
- 40 J. G. Vos, T. A. Wezendonk, A. W. Jeremiasse and M. T. M. Koper, *J. Am. Chem. Soc.*, 2018, **140**, 10270–10281.
- 41 T. Moehl, J. Suh, L. Sévery, R. Wick-Joliat and S. D. Tilley, *ACS Appl. Mater. Interfaces*, 2017, **9**, 43614–43622.
- 42 P. Poodt, D. C. Cameron, E. Dickey, S. M. George, V. Kuznetsov, G. N. Parsons, F. Roozeboom, G. Sundaram and A. Vermeer, *J. Vac. Sci. Technol., A*, 2012, **30**, 010802.
- 43 J. Divisek, H. Schmitz and B. Steffen, *Electrochim. Acta*, 1994, **39**, 1723–1731.
- 44 M. Grdeń, M. Alsabet and G. Jerkiewicz, *ACS Appl. Mater. Interfaces*, 2012, **4**, 3012–3021.
- 45 G. Zhang, Y.-S. Feng, W.-T. Lu, D. He, C.-Y. Wang, Y.-K. Li, X.-Y. Wang and F.-F. Cao, *ACS Catal.*, 2018, **8**, 5431–5441.
- 46 T. A. Ho, C. Bae, H. Nam, E. Kim, S. Y. Lee, J. H. Park and H. Shin, *ACS Appl. Mater. Interfaces*, 2018, **10**, 12807–12815.
- 47 L. X. Chen, Z. W. Chen, Y. Wang, C. C. Yang and Q. Jiang, *ACS Catal.*, 2018, **8**, 8107–8114.
- 48 F. Lyu, Y. Bai, Z. Li, W. Xu, Q. Wang, J. Mao, L. Wang, X. Zhang and Y. Yin, *Adv. Funct. Mater.*, 2017, **27**, 1702324.
- 49 J. Luo, J. H. Im, M. T. Mayer, M. Schreier, M. K. Nazeeruddin, N. G. Park, S. D. Tilley, H. J. Fan and M. Gratzel, *Science*, 2014, **345**, 1593–1596.
- 50 R. Frydendal, E. A. Paoli, I. Chorkendorff, J. Rossmeisl and I. E. L. Stephens, *Adv. Energy Mater.*, 2015, **5**, 1500991.
- 51 S. Hu, M. H. Richter, M. F. Lichterman, J. Beardslee, T. Mayer, B. S. Brunshwig and N. S. Lewis, *J. Phys. Chem. C*, 2016, **120**, 3117–3129.
- 52 T. Sun, J. Wang, X. Chi, Y. Lin, Z. Chen, X. Ling, C. Qiu, Y. Xu, L. Song, W. Chen and C. Su, *ACS Catal.*, 2018, **8**, 7585–7592.
- 53 S. Liu, M. Li, C. Wang, P. Jiang, L. Hu and Q. Chen, *ACS Sustainable Chem. Eng.*, 2018, **6**, 9137–9144.
- 54 G. V. Gibbs, R. T. Downs, C. T. Prewitt, K. M. Rosso, N. L. Ross and D. F. Cox, *J. Phys. Chem. B*, 2005, **109**, 21788–21795.
- 55 L. L. Feng, G. Yu, Y. Wu, G. D. Li, H. Li, Y. Sun, T. Asefa, W. Chen and X. Zou, *J. Am. Chem. Soc.*, 2015, **137**, 14023–14026.
- 56 J. H. Wang, Z. Cheng, J. L. Bredas and M. Liu, *J. Chem. Phys.*, 2007, **127**, 214705.
- 57 G. A. Saunders, *Contemp. Phys.*, 1973, **14**, 149–166.
- 58 S. M. Rommel and R. Wehrich, *Chem.–Eur. J.*, 2015, **21**, 9863–9867.
- 59 J. N. Schrauben, R. Hayoun, C. N. Valdez, M. Braten, L. Fridley and J. M. Mayer, *Science*, 2012, **336**, 1298–1301.
- 60 C. N. Valdez, M. Braten, A. Soria, D. R. Gamelin and J. M. Mayer, *J. Am. Chem. Soc.*, 2013, **135**, 8492–8495.
- 61 S. Intikhab, J. D. Snyder and M. H. Tang, *ACS Catal.*, 2017, **7**, 8314–8319.
- 62 Y. R. Zheng, P. Wu, M. R. Gao, X. L. Zhang, F. Y. Gao, H. X. Ju, R. Wu, Q. Gao, R. You, W. X. Huang, S. J. Liu, S. W. Hu, J. Zhu, Z. Li and S.-H. Yu, *Nat. Commun.*, 2018, **9**, 2533.
- 63 H. Feng, Z. Xu, L. Ren, C. Liu, J. Zhuang, Z. Hu, X. Xu, J. Chen, J. Wang, W. Hao, Y. Du and S. X. Dou, *ACS Catal.*, 2018, **8**, 4288–4293.
- 64 F. Farges, G. E. Brown and J. J. Rehr, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1997, **56**, 1809–1819.
- 65 Y. Li, K. A. Kuttiyiel, L. Wu, Y. Zhu, E. Fujita, R. R. Adzic and K. Sasaki, *ChemSusChem*, 2017, **10**, 68–73.
- 66 I. Harada, T. Suzuki and A. Kotani, *J. Phys. Soc. Jpn.*, 1996, **65**, 3016–3020.