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Efficient separation of C₂ hydrocarbons in a permanently porous hydrogen-bonded organic framework†

Tae-Ung Yoon,^{‡a} Seung Bin Baek,^{ib} ‡*^b Dongwook Kim,^c Eun-Jung Kim,^a Wang-Geun Lee,^{ib} Bhupendra Kumar Singh,^b Myoung Soo Lah,^{ib} ^c Youn-Sang Bae*^a and Kwang S. Kim^{ib} *^b

A highly robust porous hydrogen-bonded organic framework (HOF) constructed by 4,4',4''-benzene-1,3,5-triyl-tris(benzoic acid) not only achieves the highest uptakes of ethylene and ethane among the HOF materials, but also exhibits unusual adsorption selectivity of C₂H₆ over other C₂ gases. Besides, it exhibits the second highest acetylene uptake among all the reported HOF materials.

Hydrogen-bonded organic frameworks (HOFs), also referred to as supramolecular organic frameworks (SOFs), are promising porous organic crystalline materials that are self-assembled through multiple hydrogen bonding interactions and other non-covalent interactions like π - π stacking.¹ Over the past few years, HOFs have attracted increasing interests due to their potential applications in several areas including hydrocarbon separation,² CO₂ separation,³ proton conduction,⁴ molecular capture⁵ and sensing.⁶ In addition, compared with other porous materials, they show inherent advantages such as solution processibility, easy purification, and reproducibility by simple recrystallization. However, the multiple non-covalent interactions generally do not guarantee the stability of a porous framework, thus it is difficult to maintain a permanent porosity upon solvent removal. Consequently, we still have witnessed a few permanently porous HOFs to date.

Carboxylic acids play an important role in directing various supramolecular structures and C₃-symmetric planar carboxylic acids are of interest because they can form hexagonal networks with voids through the formation of carboxylic acid dimers.^{1,7} In fact, Hisaki and coworkers demonstrated that hexagonal

networks prepared with various C₃-symmetric planar carboxylic acids can be used as structural motifs for porous HOFs.⁸ Similarly, Zentner and coworkers recently reported a microporous HOF of C₃-symmetric 4,4',4''-benzene-1,3,5-triyl-tris(benzoic acid) (H₃BTB) *via* a hexagonal (6,3) honeycomb structural motif (Fig. 1a) in ethanol and propanols.⁹ The π - π stacking and eight-fold interpenetration of two-dimensional (2D) hexagonal sheets in this HOF-BTB generates one-dimensional (1D) solvent channel in the framework (Fig. 1b). Meanwhile, H₃BTB is also known to form nonporous DMF solvate crystals (H₃BTB-DMF) in *N,N*-dimethylformamide (DMF) where the carboxylic acids are masked by DMF molecules, prohibiting the formation of carboxylic acid dimers.¹⁰

Herein, we report single-component adsorption properties of light hydrocarbons (CH₄, C₂H₂, C₂H₄, and C₂H₆) in a robust microporous hydrogen-bonded organic framework HOF-BTB prepared from the H₃BTB-DMF crystals in methanol. We also report efficient selective separations of C₂/CH₄ binary gas mixtures in HOF-BTB. In our effort to find new self-assembled H₃BTB structures, we found that H₃BTB-DMF crystals readily recrystallize in methanol to

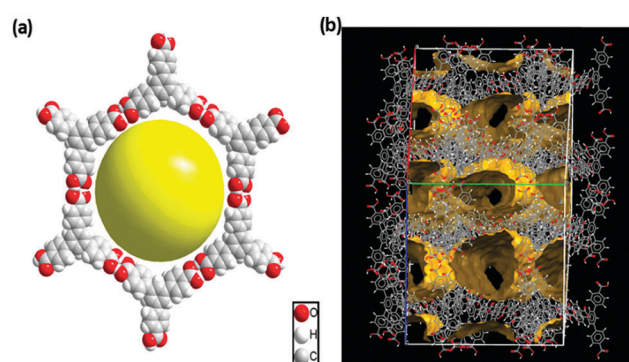


Fig. 1 (a) 2D Hexagonal sheet with a unimodal 3-c hcb (honeycomb) topology in HOF-BTB formed through carboxylic acid dimers. The yellow ball represents a one-dimensional channel with a diameter of ca. 25 nm. (b) Solvent-accessible void space of HOF-BTB calculated with the probe radius of 1.4 Å.

^a Department of Chemical and Biomolecular Engineering, and Graduate School of Integrated Engineering, Yonsei University, 50 Yonsei-ro, Seodaemun-gu, Seoul 03722, Korea. E-mail: mowbae@yonsei.ac.kr

^b Center for Superfunctional Materials and Department of Chemistry, Ulsan National Institute of Science and Technology (UNIST), 50 UNIST-gil, Ulsan 44919, Korea. E-mail: sepbirth@gmail.com, kimks@unist.ac.kr

^c Department of Chemistry, Ulsan National Institute of Science and Technology (UNIST), 50 UNIST-gil, Ulsan 44919, Korea

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‡ These authors contributed equally to this work.

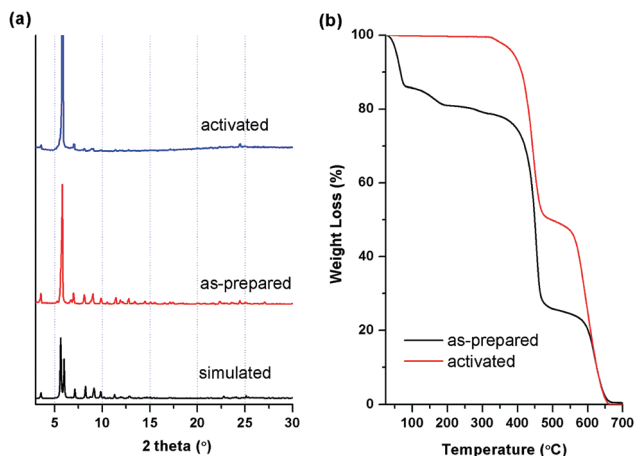


Fig. 2 (a) Powder X-ray diffraction patterns and (b) TGA curves of as-prepared and activated HOF-BTB crystals.

produce clear, transparent crystals suitable for single crystal X-ray diffraction. H_3BTB is practically insoluble in methanol so that it is inefficient to attempt direct crystallization of H_3BTB in methanol. Instead, our finding shows that soaking H_3BTB -DMF crystals in methanol at 50 °C facilitates recrystallization of H_3BTB -DMF crystals and the resulting crystals were found to exhibit essentially the same structure as the previously reported HOF-BTB by Zentner.⁹ The bulk purity of HOF-BTB crystals is examined by powder X-ray diffraction (PXRD). An observed PXRD pattern of as-prepared HOF-BTB crystals is in good agreement to the simulated PXRD pattern obtained from the reported single crystal structure (Fig. 2a). The crystals of HOF-BTB were treated under high vacuum at 120 °C for 15 hours to generate an activated sample. The PXRD pattern of activated HOF-BTB crystals is identical to that of as-prepared HOF-BTB crystals, which means that the HOF-BTB crystals are very stable even after the removal of solvent molecules. Furthermore, we also found that the immersion of the activated crystals in water for 20 days has no effect on structural stability of the activated HOF-BTB crystals (Fig. S1, ESI†). Thermogravimetric analysis (TGA) curve indicates that HOF-BTB can be stable over 300 °C, resulted from the strong hydrogen bonding interactions as well as the consecutive strong π - π stacking interactions. In addition, the TGA curve of the activated HOF-BTB sample reveals that no residual solvent molecules remain in the activated HOF-BTB (Fig. 2b).

The high stability and the large solvent-accessible void space of HOF-BTB allowed us to investigate its permanent porosity. The N_2 adsorption isotherm of the activated sample obtained at 77 K shows rapid N_2 uptake at relative low pressures ($P/P_0 < 0.1$), which is typical for a Type I isotherm curve and is also typical characteristic for microporous solids (Fig. S2a, ESI†). The total uptake of N_2 at relative pressure $P/P_0 = 1$ is $256 \text{ cm}^3 \text{ g}^{-1}$. The BET and Langmuir surface areas are estimated to be 955 and $1128 \text{ m}^2 \text{ g}^{-1}$, respectively. The mean pore diameter and total pore volume are determined to be 16.6 Å and $0.40 \text{ cm}^3 \text{ g}^{-1}$, respectively. In addition, we calculated micropore size distribution using nonlocal density functional theory (NLDFT) method (Fig. S2b, ESI†). HOF-BTB has a very narrow micropore size distribution of 1.1 to 1.5 nm diameter and a distribution peak at

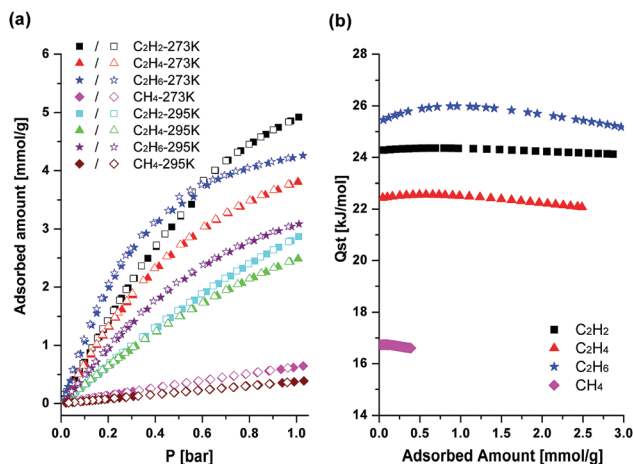


Fig. 3 (a) The adsorption isotherms of HOF-BTB for C_2H_2 , C_2H_4 , C_2H_6 and CH_4 at 273 K and 295 K, respectively, and (b) loading dependent isosteric heat (Q_{st}) of adsorption. The Q_{st} values were calculated based on adsorption isotherms at 273 K and 295 K by using the Virial method.

1.2 nm. The high surface area and total pore volume of HOF-BTB indicates that HOF-BTB possesses a highly porous structure. In order to evaluate the adsorption properties for small hydrocarbons, we performed the adsorption experiments for CH_4 , C_2H_2 , C_2H_4 , and C_2H_6 at 273 K and 295 K, respectively. Fig. 3a shows their adsorption isotherms at both temperatures. All adsorption isotherms are fully reversible in HOF-BTB, indicating pure physisorption processes and easy regeneration of HOF-BTB after gas adsorption. At 273 K, HOF-BTB is the most effective adsorbent for C_2H_2 , taking up C_2H_2 up to $110.3 \text{ cm}^3 \text{ g}^{-1}$ (4.92 mmol g^{-1}) at 1 bar. The uptake capacities of HOF-BTB for C_2H_4 and C_2H_6 are also fairly high at 273 K and 1 bar, with the values of $85.3 \text{ cm}^3 \text{ g}^{-1}$ (3.80 mmol g^{-1}) and $95.4 \text{ cm}^3 \text{ g}^{-1}$ (4.25 mmol g^{-1}), respectively. Meanwhile, at 295 K and 1 bar, HOF-BTB shows a higher C_2H_6 uptake ($69.2 \text{ cm}^3 \text{ g}^{-1}$, 3.09 mmol g^{-1}) than C_2H_2 ($64.3 \text{ cm}^3 \text{ g}^{-1}$, 2.87 mmol g^{-1}) and C_2H_4 ($55.7 \text{ cm}^3 \text{ g}^{-1}$, 2.48 mmol g^{-1}) uptakes. Among the C_2 hydrocarbons, C_2H_4 has the lowest adsorption capacities at both temperatures. Interestingly, at 273 K, HOF-BTB adsorbs more of C_2H_6 than the others at lower pressures. As the pressure increases, the adsorption amount of C_2H_2 exceeds the adsorption amount of C_2H_6 . At 295 K, on the other hand, the adsorption amount of C_2H_6 exceeds the adsorption amount of C_2H_2 and C_2H_4 in the entire pressure range. This is remarkable because selective adsorption of C_2H_6 over C_2H_2 and C_2H_4 is rarely observed in porous materials including MOFs and zeolites.¹¹ The higher affinity for C_2H_6 , the largest among C_2 hydrocarbons, might suggest that the molecular sizes (C_2H_2 , $3.32 \times 3.34 \times 5.70 \text{ Å}$; C_2H_4 , $3.28 \times 4.18 \times 4.84 \text{ Å}$; C_2H_6 , $3.81 \times 4.08 \times 4.82 \text{ Å}$)^{2e} and/or kinetic diameters (C_2H_2 , 3.3 Å; C_2H_4 , 4.2 Å; C_2H_6 , 4.4 Å)¹² of the C_2 hydrocarbons are less important in the adsorption of C_2 hydrocarbons. It might be likely that van der Waals interactions between C_2 hydrocarbons and HOF-BTB are more crucial factor for C_2 hydrocarbons adsorption than the molecular sizes of the gases at lower pressures. The van der Waals interaction is generally stronger as the molecular

Table 1 Comparison of C₂H₂, C₂H₄, and C₂H₆ uptakes in reported HOFs

| HOFs | S _{ABET} (m ² g ⁻¹) | C ₂ H ₂ (cm ³ g ⁻¹) | | C ₂ H ₄ (cm ³ g ⁻¹) | | C ₂ H ₆ (cm ³ g ⁻¹) | | Q _{st} (kJ mol ⁻¹) | | | Ref. |
|---------|---|--|-------|--|-------|--|-------|---|-------------------------------|-------------------------------|-----------|
| | | 273 K | 295 K | 273 K | 295 K | 273 K | 295 K | C ₂ H ₂ | C ₂ H ₄ | C ₂ H ₆ | |
| SOF-1a | 474 | 61 | 50 | — | — | — | — | 36.2 | — | — | 2f |
| HOF-1a | 359 | 63 | 57 | 8.3 | 3.9 | — | — | 58.1 | 31.9 | — | 2e |
| HOF-3a | 165 | 58 | 47 | — | — | — | — | 18 | — | — | 2c |
| HOF-4a | 312 | — | — | 17.3 | 11.1 | 5.1 | 3.6 | — | 44 | 14 | 2d |
| HOF-5a | 1101 | 182 | 102 | — | — | — | — | 27.6 | — | — | 3c |
| HOF-BTB | 955 | 110.3 | 64.3 | 85.3 | 55.7 | 95.4 | 69.2 | 24.3 | 22.6 | 25.4 | This work |

size increases, so it can be expected that the largest molecule C₂H₆ is the most efficient to interact with HOF-BTB. It should be noted that HOF-BTB achieves the highest C₂H₄ and C₂H₆ uptakes among all the HOF materials reported to date (Table 1). Compared with covalent organic frameworks (COFs), C₂H₆ uptake of HOF-BTB is comparable to most reported COFs. ZnP-CTF-500^{13a} and T-COF^{13b} outperform the C₂H₆ uptake capacity of HOF-BTB (Table S1, ESI[†]). Furthermore, C₂H₄ uptake capacity of HOF-BTB surpasses those of the reported COFs such as DBA-3D-COF-1^{13c} and MCOF-1.^{13d} It also exhibits the second highest C₂H₂ uptake among the HOF materials reported to date. Only HOF-5a exceeds the C₂H₂ uptake amount of HOF-BTB. In contrast to the high C₂ hydrocarbons uptakes, much less amount of CH₄, with the values of 14.5 and 8.7 cm³ g⁻¹, is adsorbed in HOF-BTB at 273 K and 295 K, respectively. The loading dependent isosteric heats (Q_{st}) of adsorption of C₂ hydrocarbons are larger than that of CH₄, which indicates that the stronger interactions between the C₂ hydrocarbons and the pore surfaces of HOF-BTB than that of CH₄ and HOF-BTB (Fig. 3b). For all the hydrocarbons, the values of Q_{st} remain relatively constant through the whole loading ranges. High affinity for C₂H₆ is also evident in the Q_{st} value. As shown in Fig. 3b, the Q_{st} value of C₂H₆ is higher than those of C₂H₂ and C₂H₄. Compared with other HOF materials, zero-coverage isosteric heats of adsorption for C₂H₂ and C₂H₄ are smaller than those of reported HOFs (Table 1). These lower values of Q_{st} suggest HOF-BTB would be a promising candidate for energy-saving separation of C₂ hydrocarbons/CH₄ mixtures.

To analyse the binary gas mixture separation abilities, the pure single-component isotherms were fitted with the Langmuir-Freundlich isotherm model (Fig. S3–S6, ESI[†]).¹⁴ Then, ideal adsorbed solution theory (IAST)¹⁵ was used to predict the separation selectivity of binary gas mixtures (50 : 50) at 273 K and 295 K (Tables S2 and S3, ESI[†]). The IAST predicts that HOF-BTB would allow selective separations of C₂ hydrocarbons over CH₄. At both temperatures, C₂H₆/CH₄ gas mixture shows the highest selectivity values among the C₂ hydrocarbon/CH₄ gas mixtures. The observed C₂H₆/CH₄ selectivity at 295 K (14) is comparable with the IAST-predicted C₂H₆/CH₄ selectivities for N-COF (18), P-COF (12) and T-COF (10) at 298 K.^{13b} The IAST also predicts the fairly good separation selectivities for C₂H₂/CH₄ and C₂H₄/CH₄ binary gas mixtures (Fig. 4a and b). Although the IAST can predict the mixture behaviours in terms of adsorption equilibrium, it cannot reflect the mixture behaviours in terms of adsorption kinetics, which cannot be ignored in real adsorption separation processes. Hence, we performed dynamic breakthrough experiments to evaluate the potential of

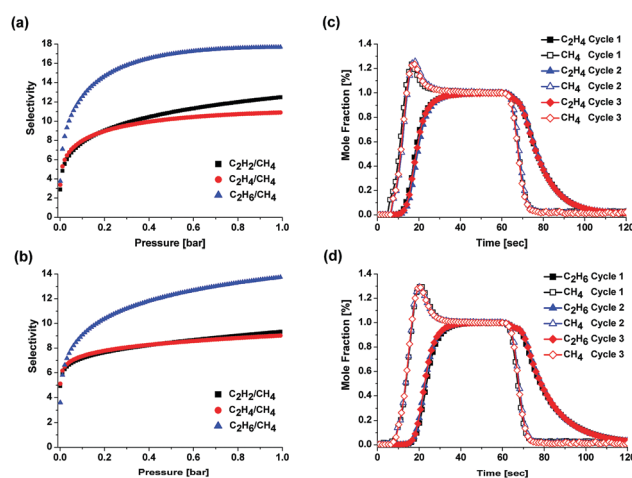


Fig. 4 The IAST-predicted binary gas mixture separation selectivities of HOF-BTB at (a) 273 K and (b) 295 K. The experimental breakthrough curves at three consecutive cycles for a packed-bed filled with HOF-BTB pellets at 298 K and total hydrocarbon pressures of 0.5 bar: (c) equimolar C₂H₄/CH₄ mixture (20 ml min⁻¹); (d) equimolar C₂H₆/CH₄ mixture (20 ml min⁻¹). For each run, 20 ml min⁻¹ of He flow was mixed with equimolar hydrocarbon mixture.

HOF-BTB for the adsorptive separation of C₂H₆/CH₄ and C₂H₄/CH₄ mixtures under dynamic flow conditions.¹⁶ Unfortunately, we could not perform the dynamic breakthrough experiment for C₂H₂/CH₄ mixture due to the University's safety regulations. Fig. 4c and d show the experimental breakthrough curves of C₂H₄/CH₄ and C₂H₆/CH₄ mixtures on a column packed with HOF-BTB pellets. In the case of C₂H₆/CH₄ separation, CH₄ elutes first from the column, whereas C₂H₆ is strongly retained. This indicates that under mixture flow condition, HOF-BTB primarily adsorbs C₂H₆ rather than CH₄. Similarly, the breakthrough curves for C₂H₄/CH₄ mixture exhibit selective C₂H₄ adsorption over CH₄. For all cases, breakthrough curves of weaker adsorbate CH₄ show “roll-up” behaviours, which are normally explained as a result of the competitive adsorption between two adsorbate species. Although the weaker adsorbate molecules initially occupy the adsorption sites in the column, some of the adsorbed molecules are replaced by stronger adsorbate molecules.¹⁷ These experimental breakthrough results indicate that HOF-BTB can effectively separate C₂H₄/CH₄ and C₂H₆/CH₄ mixtures under dynamic flow conditions. Easy regeneration of the used adsorbent is an essential consideration for practical operation of an adsorptive separation process. After performing a breakthrough experiment, we regenerated

the column by simply purging it for 10 min under a He flow of 40 ml min⁻¹ without heating the column. As displayed in Fig. 4c and d, essentially identical breakthrough curves were produced during the three consecutive cycles for both gas mixtures. This is remarkable because the regenerations were performed under mild conditions. This mild regeneration condition could be attributed to the relatively small Q_{st} values for C₂ hydrocarbons.

In summary, we have reported a highly robust microporous hydrogen-bonded organic framework HOF-BTB, which shows the higher adsorption capacities for C₂ hydrocarbons at 273 K and 295 K. The adsorption amounts of C₂H₄ and C₂H₆ reach the highest values among all the reported HOFs, while C₂H₂ uptake amount reaches the second highest values among all the HOFs. Breakthrough experiments reveal that HOF-BTB can selectively separate C₂H₄ and C₂H₆ from CH₄ under dynamic mixture flow conditions.

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Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 Y.-F. Han, Y.-X. Yuan and H.-B. Wang, *Molecules*, 2017, **22**, 266.
- 2 (a) A. Pulido, L. Chen, T. Kaczorowski, D. Holden, M. A. Little, S. Y. Chong, B. J. Slater, D. P. McMahon, B. Bonillo, C. J. Stackhouse, A. Stephenson, C. M. Kane, R. Clowes, T. Hasell, A. I. Cooper and G. M. Day, *Nature*, 2017, **543**, 657; (b) F. Hu, C. Liu, M. Wu, J. Pang, F. Jiang, D. Yuan and M. Hong, *Angew. Chem., Int. Ed.*, 2017, **56**, 2101; (c) P. Li, Y. He, Y. Zhao, L. Weng, H. Wang, R. Krishna, H. Wu, W. Zhou, M. O'Keefe, Y. Han and B. Chen, *Angew. Chem., Int. Ed.*, 2015, **54**, 574; (d) P. Li, Y. He, H. D. Arman, R. Krishna, H. Wang, L. Weng and B. Chen, *Chem. Commun.*, 2014, **50**, 13081; (e) Y. He, S. Xiang and B. Chen, *J. Am. Chem. Soc.*, 2011, **133**, 14570; (f) W. Yang, A. Greenaway, X. Lin, R. Matsuda, A. J. Blake, C. Wilson, W. Lewis, P. Hubberstey, S. Kitagawa, N. R. Champness and M. Schröder, *J. Am. Chem. Soc.*, 2010, **132**, 14457.
- 3 (a) R. S. Patil, D. Banerjee, C. Zhang, P. K. Thallapally and J. L. Atwood, *Angew. Chem., Int. Ed.*, 2016, **55**, 4523; (b) S. Nandi, D. Chakraborty and R. Vaidhyanathan, *Chem. Commun.*, 2016, **52**, 7249; (c) H. Wang, B. Li, H. Wu, T.-L. Hu, Z. Yao, W. Zhou, S. Xiang and B. Chen, *J. Am. Chem. Soc.*, 2015, **137**, 9963; (d) W. Yang, B. Li, H. Wang, O. Alduhaish, K. Alfooty, M. A. Zayed, P. Li, H. D. Arman and B. Chen, *Cryst. Growth Des.*, 2015, **15**, 2000; (e) J. Lü, C. Perez-Krap, M. Suyetin, N. H. Alsmail, Y. Yan, S. Yang, W. Lewis, E. Bichoutskaia, C. C. Tang, A. J. Blake, R. Cao and M. Schröder, *J. Am. Chem. Soc.*, 2014, **136**, 12828; (f) X.-Z. Luo, X.-J. Jia, J.-H. Deng, J.-L. Zhong, H.-J. Liu, K.-J. Wang and D.-C. Zhong, *J. Am. Chem. Soc.*, 2013, **135**, 11684; (g) M. Mastalerz and I. M. Oppel, *Angew. Chem., Int. Ed.*, 2012, **51**, 5252.
- 4 (a) A. Karmakar, R. Illathalappil, B. Anothumakkool, A. Sen, P. Samanta, A. V. Desai, S. Kurungot and S. K. Ghosh, *Angew. Chem., Int. Ed.*, 2016, **55**, 10667; (b) W. Yang, F. Yang, T.-L. Hu, S. C. King, H. Wang, H. Wu, W. Zhou, J.-R. Li, H. D. Arman and B. Chen, *Cryst. Growth Des.*, 2016, **16**, 5831.
- 5 (a) W. Yan, X. Yu, T. Yan, D. Wu, E. Ning, Y. Qi, Y.-F. Han and Q. Li, *Chem. Commun.*, 2017, **53**, 3677; (b) T.-H. Chen, I. Popov, W. Kaveevivitchai, Y.-C. Chuang, Y.-S. Chen, O. Daugulis, A. J. Jacobson and O. S. Miljanić, *Nat. Commun.*, 2014, **5**, 5131.
- 6 (a) H. Wang, H. Wu, J. Kan, G. Chang, Z. Yao, B. Li, W. Zhou, S. Xiang, J. C.-G. Zhao and B. Chen, *J. Mater. Chem. A*, 2017, **5**, 8292; (b) Z. Sun, Y. Li, L. Chen, X. Jing and Z. Xie, *Cryst. Growth Des.*, 2015, **15**, 542; (c) H. Zhou, Q. Ye, X. Wu, J. Song, C. M. Cho, Y. Zong, B. Z. Tang, T. S. Andy Hor, E. K. Lee Yeow and J. Xu, *J. Mater. Chem. C*, 2015, **3**, 11874.
- 7 O. Ivasenko and D. F. Perepichka, *Chem. Soc. Rev.*, 2011, **40**, 191.
- 8 (a) I. Hisaki, S. Nakagawa, N. Ikenaka, Y. Imamura, M. Katouda, M. Tashiro, H. Tsuchida, T. Ogoshi, H. Sato, N. Tohnai and M. Miyata, *J. Am. Chem. Soc.*, 2016, **138**, 6671; (b) I. Hisaki, S. Nakagawa, N. Tohnai and M. Miyata, *Angew. Chem., Int. Ed.*, 2015, **54**, 3008.
- 9 C. A. Zentner, H. W. H. Lai, J. T. Greenfield, R. A. Wisconsin, M. Zeller, C. F. Campana, O. Talu, S. A. FitzGerald and J. L. C. Rowsell, *Chem. Commun.*, 2015, **51**, 11642.
- 10 (a) S. B. Baek, D. Moon, R. Graf, W. J. Cho, S. W. Park, T.-U. Yoon, S. J. Cho, I.-C. Hwang, Y.-S. Bae, H. W. Spiess, H. C. Lee and K. S. Kim, *Proc. Natl. Acad. Sci. U. S. A.*, 2015, **112**, 14156; (b) E. Weber, M. Hecker, E. Koeppe, W. Orlia, M. Czugler and I. Csöreg, *J. Chem. Soc., Perkin Trans. 2*, 1988, 1251.
- 11 (a) P.-Q. Liao, W.-X. Zhang, J.-P. Zhang and X.-M. Chen, *Nat. Commun.*, 2015, **6**, 8697; (b) Y. Chen, H. Wu, D. Lv, R. Shi, Y. Chen, Q. Xia and Z. Li, *Ind. Eng. Chem. Res.*, 2018, **57**, 4063.
- 12 J.-R. Li, R. J. Kuppler and H.-C. Zhou, *Chem. Soc. Rev.*, 2009, **38**, 1477–1504.
- 13 (a) H. Ma, H. Ren, S. Meng, F. Sun and G. Zhu, *Sci. Rep.*, 2013, **3**, 2611; (b) J. Dong, Y. Wang, G. Liu, Y. Cheng and D. Zhao, *CrystEngComm*, 2017, **19**, 4899; (c) L. A. Baldwin, J. W. Crowe, D. A. Pyles and P. L. McGrier, *J. Am. Chem. Soc.*, 2016, **138**, 15134; (d) H. Ma, H. Ren, S. Meng, Z. Yan, H. Zhao, F. Sun and G. Zhu, *Chem. Commun.*, 2013, **49**, 9773.
- 14 J. A. Mason, K. Sumida, Z. R. Herm, R. Krishna and J. R. Long, *Energy Environ. Sci.*, 2011, **4**, 3030.
- 15 A. L. Myers and J. M. Prausnitz, *AIChE J.*, 1965, **11**, 121.
- 16 S.-J. Lee, K. C. Kim, T.-U. Yoon, M.-B. Kim and Y.-S. Bae, *Microporous Mesoporous Mater.*, 2016, **236**, 284–291.
- 17 R. T. Yang, *Adsorbents: Fundamentals and Applications*, John Wiley & Sons, 2003.