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Crosslinked polymeric bis(triphenylphosphine) iminium chloride gels as recyclable catalysts

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Bis(triphenylphosphine)iminium chloride, PPNCl, is an increasingly important catalyst due to its broad reactivity in both small molecule and polymer synthesis. A recently reported polymeric variant, poly(PPNCl), offers limited recyclability in small molecule synthesis, but its soluble, thermoplastic nature complicates its recovery and reuse. To address this limitation, we developed a crosslinked poly(PPNCl) thermoset that dramatically improves performance and reaction scope. As a catalyst in CO₂/epoxide coupling reactions the crosslinked derivative is faster and more readily recycled. In addition, the reaction scope can be extended to polymer synthesis, with soluble polyesters formed from coupling epoxides with maleic anhydrides isolated by filtration, with activity retained over multiple reaction cycles. Remarkably, the reaction scope can be further extended by anion exchange, with crosslinked poly(PPNCo(CO)₄) used to catalyse the carbonylative ring expansion of epoxides to β -lactones. Together these show that crosslinked poly(PPNCl) is a promising platform for metal-free, robust, recyclable catalysis.

Introduction

Catalysts are essential enablers of sustainable chemical synthesis. Metal-free catalysts are of growing importance, with reduced cost, footprint and toxicity sparking significant academic and industrial interest.^{1–4} Bis(triphenylphosphine)iminium (PPN) salts are a flourishing sub-class of catalysts, particularly notable for their excellent solubility, stability, reaction scope and ease of use.^{5–9} The chloride derivative, PPNCl, has become a valuable catalyst, particularly for CO₂ coupling to form small molecules^{10–12} and synthesis of degradable polyesters from the copolymerisation of cyclic anhydrides and epoxides.^{13–17}

However, the economic and environmental cost of the complex PPN cation limits sustainability assurances. Previous work in our group reported a polymeric bis(triphenylphosphine)iminium chloride (poly(PPNCl)) variant as a catalyst for multiple small molecule transformations; precipitation allowed for catalyst recovery and reuse (Fig. 1a).¹⁸ Our efforts to extend the use of poly(PPNCl) to polymer synthesis, such as phthalic anhydride and cyclohexene oxide copolymerisation, highlighted significant limitations, with their analogous solubilities preventing separation and recycling. The use of dissolution and precipitation consistently led to product loss,

meaning that even with good activity retention the system lacked long term sustainability.

Crosslinked polymer-supported catalysts, with the active catalytic species immobilized onto a crosslinked polymer matrix, are solid supports reacting with gases or liquids.^{19–21} Compared to soluble polymer-supported catalysts, they offer easy separation from the reaction mixture through filtration or in flow, offering future opportunities to manufacture in large-

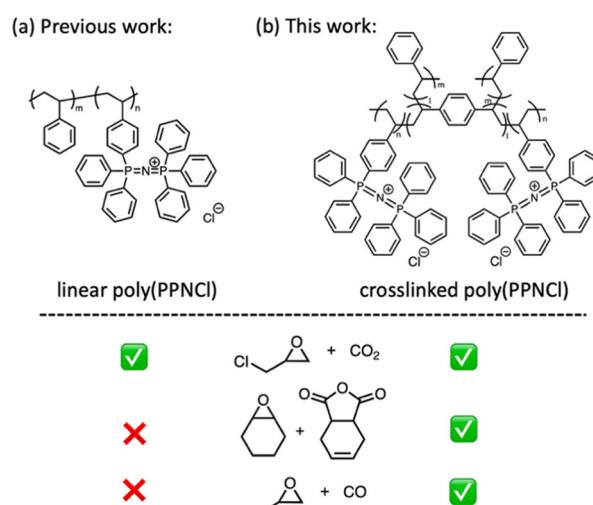


Fig. 1 Comparing previously reported linear poly(PPNCl) with cross-linked poly(PPNCl) systems enabling recyclable catalysis of epoxides reacting through CO₂ insertions, copolymerisations with anhydrides, and CO insertions.

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scale reactors and with continuous production.^{22–24} These catalysts can offer improved thermal and mechanical stability and anti-solvent free recyclability while maintaining good catalytic efficiency.²⁵ Owing to these advantages, they have been explored widely, in diverse reactions, suggesting a crucial role in future sustainable manufacturing processes.^{26–28} Together, this suggested the need to shift topology, framing the new target of a crosslinked poly(PPNCl) that would swell but not dissolve to facilitate recovery.

In this work, we report the preparation of a crosslinked poly(PPNCl) (Fig. 1b). A comparison of the catalytic reactivity between soluble linear catalysts and crosslinked poly(PPNCl) gels showed increased catalytic activity and improved recyclability for both small molecule and, for the first time, polymeric syntheses. We further used the robustness of these gels to support anion exchange to synthesise a poly(PPNCo(CO)₄) variant, inspired by a rare report of a small molecule derivative functioning in CO coupling reactions in conjunction with a Lewis acid.²⁹ Indeed, we were able to show that crosslinked poly(PPNCo(CO)₄) is an effective catalyst for the carbonylative ring expansion of epoxides to β -lactones.

Results and discussion

Synthesis of crosslinked poly(PPNCl)

A crosslinked precursor, **P1**, was formed *via* radical copolymerisation of styrene, St, *p*-(diphenylphosphino)styrene, StPPh₂ and divinyl benzene (DVB), as shown in Scheme 1a. After optimisation, a feed ratio of 95 : 5 : 1 of St, StPPh₂ and DVB afforded a stiff gel. ³¹P NMR spectroscopy of the swollen gel still afforded good resolution, with the crude product showing a narrow resonance at -5.8 ppm and a broad resonance at -6.1 ppm, associated with unpolymerised StPPh₂ monomers

and the incorporated PPh₃ anchors, respectively (Fig. S1). Conversion of StPPh₂ was 98% from integrating these resonances. Immersion of **P1** in several aliquots of toluene for 4 h facilitated the extraction of residual monomers, offering pure **P1** which was subsequently dried *in vacuo* and manually ground into a fine powder.

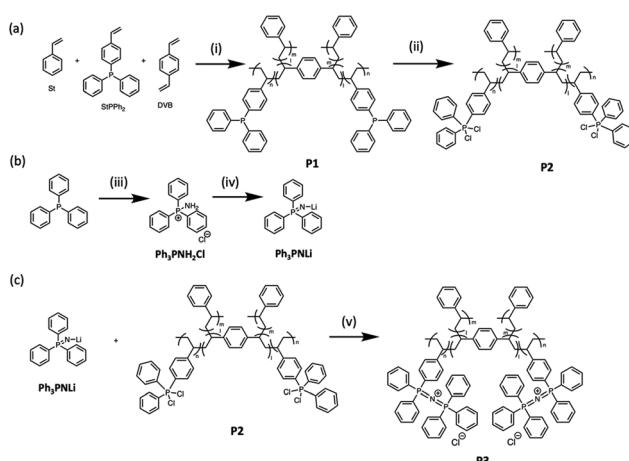
Post-polymerisation of this thermoset polymer was accomplished by modifying the methodology used to prepare poly(PPNCl).¹⁸ Dichlorination of **P1** with C₂Cl₆ accessed the phosphorus(v) StPPh₂Cl₂ crosslinked gel **P2** (Scheme 1a). A broad resonance at 63.5 ppm was observed in the ³¹P NMR spectrum (Fig. S2), comparable to that of its small molecule analogue (65.0 ppm). No signal associated with **P1** was observed, suggesting full conversion. To remove any remaining C₂Cl₆, the crude product was immersed in aliquots of THF for 4 h, the solvent was decanted, and the product was dried *in vacuo* and manually ground into a fine powder.

Finally, the targeted crosslinked poly(PPNCl), **P3**, was synthesised by reaction between a swollen **P2** gel and Ph₃PNLi (Scheme 1c). A broad peak at 20.7 ppm was attributed to the crosslinked poly(PPNCl) functionalities (Fig. S3), with the disappearance of the **P2** resonance again suggesting full conversion. Small signals at 29.3 ppm, suggesting phosphine oxidation, and 35.6 ppm, attributed to the residual Ph₃PNH₂Cl (35.6 ppm), were observed. Purification of **P3** by dialysis removed the majority of small molecule impurities; however *ca.* 6% of partially oxidised phosphine remained in the product (Fig. S4). As this phosphine oxide was not expected to interfere with catalyst performance, **P3** was used without further purification. Importantly, no progression in the extent of oxidation was observed through catalyst use or reuse.

CO₂/GC coupling reaction

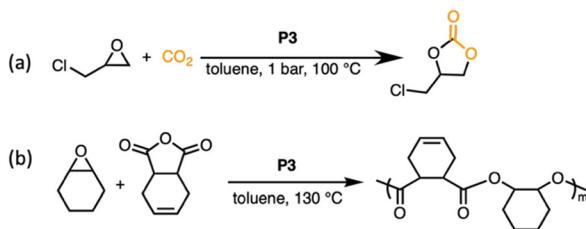
Crosslinked catalysts often depend on swelling to promote catalytic activity, improving the accessibility of reactants to active sites within the polymer network. While linear poly(PPNCl) showed good catalytic reactivity for the insertion of CO₂ to epoxides,¹⁸ we wished to assess the comparative catalytic reactivity of the crosslinked variant, **P3**, in the coupling reaction between CO₂ and glycidyl chloride (GC) to yield cyclic carbonates (Scheme 2).

Under previously optimised conditions using linear poly(PPNCl), the conversion of GC to chloropropene carbonate reached up to 98% in neat reactions.¹⁸ With the same % functional loading (0.4%), **P3** was employed for the CO₂/GC coupling reaction. Given that the buried catalytic sites of crosslinked poly(PPNCl) become more accessible in the swollen state, toluene was introduced to assist in swelling the gel, despite decreasing effective catalyst concentration. The results are summarized in Table S1. As a control experiment, the reaction catalysed by linear poly(PPNCl) in toluene gave 97% conversion of GC (entry 2, Table S1). Notably, when switching to crosslinked poly(PPNCl), full conversion of GC to chloropropene carbonate was achieved (entry 3, Table S1), displaying comparable catalytic reactivity to that of its linear analogue poly(PPNCl). The porous structure of the swollen network



Scheme 1 (a) Synthetic procedures of the crosslinked precursor **P2** from St, StPPh₂ and the crosslinker DVB with a feed ratio of 95 : 5 : 1; (i) AIBN, toluene, 70 °C, 16 h and (ii) C₂Cl₆, THF, r.t., 16 h. (b) Synthesis of Ph₃PNLi from triphenylphosphine; (iii) hydroxylamine-O-sulfonic acid, methanol/DCM, r.t., 1 h; (iv) *n*-butyllithium, THF, initiated at -50 °C -15 °C 2 h. (c) Synthesis of **P3**; (v) THF, r.t., 16 h.





Scheme 2 (a) CO_2/GC coupling reaction catalysed by crosslinked poly(PPNCl). General reaction conditions: reactions were carried out in a sealed flask on a 12.8 mmol scale of GC and 0.4 mol% crosslinked poly(PPNCl) (mol% as calculated from the number of PPNCl moieties) in 1 mL toluene at 100 $^{\circ}\text{C}$ and 1 bar. (b) Ring-opening polymerisation catalysed by crosslinked poly(PPNCl). General reaction conditions: polymerisation was conducted in toluene at 130 $^{\circ}\text{C}$. **CHO : CDA : PPNCl group** = 200 : 100 : 4.

allows both CO_2 and GC to diffuse through the network and readily access catalytic sites embedded within the gel. No polyether formation was observed in the ^1H NMR spectrum (Fig. S5), indicating high selectivity for cyclic carbonate formation.

Exploring the limits of this new crosslinked system for CO_2/GC coupling, we reduced the reaction times (Table S2). Only at 4 h (95%) or 2 h (73%) did conversion not reach 100%. Compared to its linear counterpart, the surprising increase in activity for the crosslinked variant suggests more than simply a retention of the number of active sites in the cross-linked system. The access of reactants to catalytic sites is potentially improved due to compartmentalisation effects that minimise diffusion distances for substrates. Solvent choice can also improve CO_2 solubility in active pockets. Indeed, reactions using the more polar anisole to swell the crosslinked gel gave lower conversions (66%), suggesting that CO_2 solubility may be an important factor.

Catalyst recyclability

Catalyst recyclability was evaluated at both 4 h and 24 h to both compare with those from linear polymer recycling studies and to assess performance below full conversion limits. Recycling studies were conducted in air. After each cycle, gels were washed with three aliquots of diethyl ether to extract cyclic carbonate products. The gels were dried *in vacuo* between catalytic runs. As illustrated in Fig. 2, the catalytic reactivity of **P3** decreases minimally over three reaction cycles (1st cycle: 95%; 2nd: 90%; 3rd: 88%) with a fixed reaction time of 4 h. No phosphine oxidation occurred after three reaction cycles as demonstrated by ^{31}P NMR spectroscopy (Fig. S6). In contrast, when the reaction time was extended to 24 h, a well-maintained conversion of GC was observed over three cycles, achieving complete transformation into chloropropene carbonate in each run. This suggests no decomposition of the catalytic sites, but reduced access to these sites in the gels after repeated drying and swelling cycles; indeed swelling appeared to be qualitatively slower in repeatedly dried systems. The recycling process was much more straightforward than those

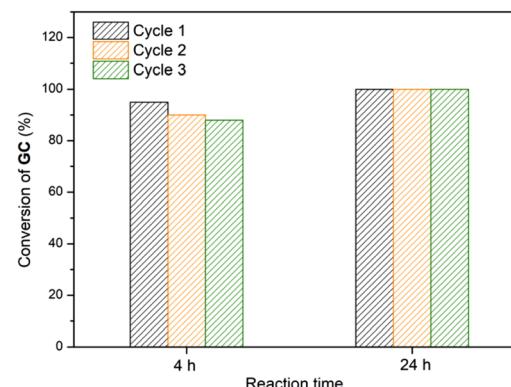


Fig. 2 Conversions of GC in the CO_2/GC coupling reaction catalysed by cross-linked poly(PPNCl) with reaction times of 4 h and 24 h over three reaction cycles. Conversion was calculated by ^1H NMR spectroscopy.

with previous systems, suggesting future application in flow may be possible. The consistency of the catalytic reactivity of crosslinked poly(PPNCl) over prolonged reaction times (24 h) supports robust stability for longer term manufacturing.

Epoxide anhydride copolymerisation

Poly(PPNCl) was unable to facilitate productive polymerisation or copolymerisation reactions due to separation challenges (Table S3). While **P3** slightly improves catalyst performance in small molecule coupling reactions, we believe that this new topology would offer an extended reaction scope. Indeed, the system exhibits high reactivity in the ring-opening copolymerisation of cyclohexene oxide (**CHO**) and *cis*-4-cyclohexene-1,2-dicarboxylic anhydride (**CDA**) (Scheme 2b and Table S3).

A near-quantitative conversion of **CHO : CDA** into the desired polymer was observed after an initial catalyst screen (24 h, entry 2, Table S3), matching previous reports of the small molecule PPNCl serving as a standalone catalyst in the ring-opening polymerisation of **CHO** and anhydrides.^{17,30} Reducing reaction times to assess catalyst performance showed that full conversion of **CDA** was achieved in as little as 2 h (Table S3). The polymerisation was monitored over the course of 2 h by ^1H NMR spectroscopy (Fig. S7). The reaction showed first order kinetics over the limited time range, with $k = 0.26 \text{ h}^{-1}$ (Fig. S8). Notably, at the latter stage of the polymerisation, when nearly all **CDA** was consumed, new resonances in the region of 3.5 ppm were observed, with these ether linkages suggesting homopolymerisation in the absence of a comonomer.

To assess the recyclability of **P3** in this copolymerisation, reaction times of 1 h and 24 h were selected (Fig. 3), with the former assessing activity below full conversion and the latter assessing long term robustness. As with small molecule transformations, catalysts were washed, dried and reswollen between catalytic runs. **CDA** conversion decreased slightly over three reaction cycles (1st cycle: 93%; 2nd: 85%; and 3rd: 80%) while consistent conversions (1st cycle: 99%; 2nd: 99%; and



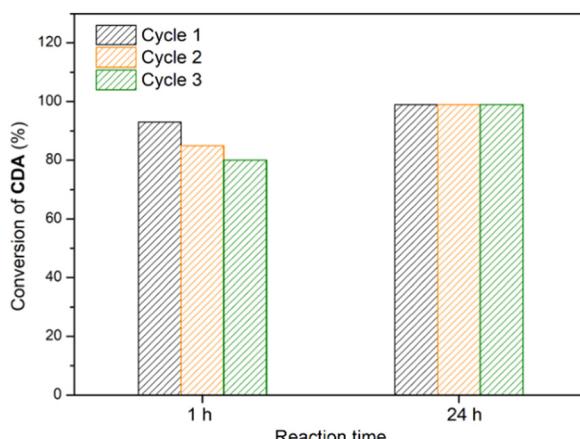


Fig. 3 Conversion of CDA in the ring-opening polymerisation of CHO and CDA catalysed by crosslinked poly(PPNCl) with reaction times of 4 h and 24 h over three reaction cycles. Conversion was calculated by ^1H NMR spectroscopy.

3rd: 99%) were observed in longer reactions, suggesting excellent recyclability and robustness.

Carbonylative ring expansion of epoxides to β -lactones

Finally, we exploited the robustness of these gels to facilitate anion exchange to further increase reaction scope. The carbonylative ring expansion of epoxides forms valuable β -lactone intermediates, and has recently attracted significant attention,^{31–33} often relying on complex bimetallic catalytic systems incorporating a Lewis acid and $\text{Co}(\text{CO})_4$ salt.^{34–36} We hypothesized that **P3** could be developed into a unique carbonylation catalyst. Thus, crosslinked poly(PPNCo(CO)₄), **P4** (Fig. 4), was prepared through an anion exchange reaction between **P3** and $\text{NaCo}(\text{CO})_4$ with conditions based on unpublished optimisation of the linear system.²⁹

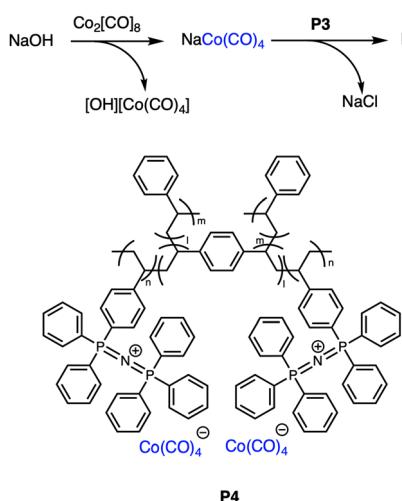


Fig. 4 A two-step synthetic method to synthesise crosslinked poly(PPNCo(CO)₄) and its structure.

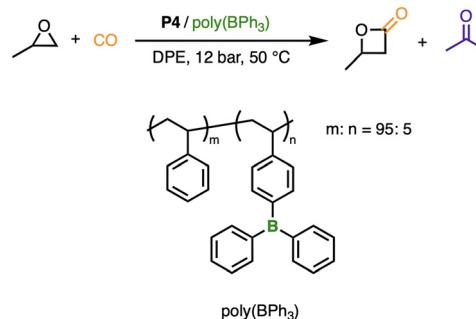


Fig. 5 Carbonylation reaction of PO catalysed by **P4**/poly(BPh₃). General reaction conditions: reactions were performed in a high-pressure reactor with 1 mol% crosslinked poly(PPNCo(CO)₄) and 1 mol% poly(BPh₃) (mol% as calculated from the number of functional groups) in 0.9 mL diphenyl ether (DPE) at 50 °C with a CO pressure of 12 bar.

Building from our previous work on polymeric frustrated Lewis pairs,^{29,37–41} we explored the catalytic activity of **P4** in conjunction with our poly(BPh₃) Lewis acid, with the carbonylation of propylene oxide (PO) as a substrate. As observed in other systems,^{36,42–44} the solvent significantly influenced both selectivity and reaction rates. Highly coordinating solvents competed with binding sites and slowed ring-opening, while non-coordinating solvents favoured polyether formation. The weakly coordinating solvent diphenyl ether, DPE, gave optimal results (Fig. 5). Importantly, the reaction was productive at much lower pressures (12 bar) than those previously reported for many systems, including a report of immobilized $\text{Co}(\text{CO})_x$ onto reduced carbon nanotubes (60 bar).⁴⁵

PO conversion to β -butyrolactone, BBL, was 77% as determined by ^1H NMR spectroscopy (Fig. S9 and Table S4), suggesting that crosslinked poly(PPNCo(CO)₄)/poly(BPh₃) is an unusually active catalyst combination. The catalytic system had high selectivity for β -lactone over polymer formation, as no polyether is observed. The formation of the off-cycle by-product acetone was observed. Varying reaction times changed selectivity for the lactone over acetone (Table S4) resulting in lower BBL yields at extended reaction times (Fig. S10 and S11). Preliminary ICP-OES measurements suggest no leaching of cobalt from catalytic reactions from the analysis of the supernatant, although more work is needed to ensure sample preparation does not lead to false negative values. Based on the above results, the crosslinked poly(PPNCo(CO)₄)/poly(BPh₃) catalytic system is a promising candidate for more sustainable carbonylation reactions and remains an active area of future research efforts.

Conclusions

In summary, crosslinked poly(PPNCl) was prepared for the first time through post-polymerisation modification. Its utility was first demonstrated in the CO_2/GC coupling reaction, where it surprisingly exhibited superior catalytic efficiency compared



to its linear analogue poly(PPNCl), producing chloropropene carbonate. Furthermore, the crosslinked polymer demonstrated excellent recyclability across multiple reaction cycles. Leveraging this robustness, crosslinked poly(PPNCl) was subsequently employed as a sole catalyst for the ring-opening polymerisation of **CHO** and **CDA**. It achieved quantitative incorporation of **CDA** into the polyester backbone within 2 hours at 130 °C using only 4 mol% catalyst. Crucially, the crosslinked catalyst could be easily separated *via* centrifugation and maintained high reactivity over three consecutive cycles. Finally, to showcase its synthetic versatility, crosslinked poly(PPNCl) was transformed into crosslinked poly(PPNCo(CO)₄). This derived catalyst successfully facilitated the formation of β -butyrolactone from **PO** and CO under mild conditions (50 °C, 12 bar CO), achieving 73.7% conversion in the presence of linear poly(BPh₃). These results confirm crosslinked poly(PPNCl) as a versatile and robust metal-free catalytic platform with excellent recyclability.

Author contributions

Ziwei Xu: investigation, methodology and manuscript writing.
Michael P. Shaver: conceptualisation, manuscript writing – review and editing, supervision and funding acquisition.

Conflicts of interest

There are no conflicts to declare.

Data availability

The data supporting this article have been included as part of the supplementary information (SI). Supplementary information: experimental details, additional characterisation data, and catalyst optimisation. See DOI: <https://doi.org/10.1039/d5py01143d>.

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