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Exploring pre-catalyst scope and selectivity in Negishi coupling of aryl iodides in deep eutectic solvents

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Negishi coupling is a valuable synthetic protocol for the formation of C–C bonds. However, the use of highly reactive alkyl- or arylzinc reagents in these reactions usually requires them to be performed under strictly anhydrous conditions using Schlenk techniques. It has recently been demonstrated that using deep eutectic solvents (DES) as reaction media allows Negishi reactions to be performed without any protection from air or moisture. This paper explores pre-catalyst scope and regioselectivity in these reactions. This has led to new protocols using cheap and readily available Pd pre-catalysts, e.g. Pd₃(OAc)₆, whilst allowing coupling of some *ortho*-functionalised aryl iodides without the need for a supporting phosphine ligand. Unusual solvent-induced selectivity is demonstrated, which allows highly regioselective C–C bond formation to take place, even when multiple C–I bonds are present on the aryl iodide precursor. To further understand this, the activity and selectivity of a range of Pd pre-catalysts towards different iodoarenes has been explored, with the results helping inform the choice of catalyst for specific substrates and desired regioselectivity. The use of DES in these reactions promises to simplify their use, especially in environments where strict exclusion of air and moisture is not possible, whilst simultaneously providing access to regioselectivity that does not occur in conventional solvents under the same conditions.

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Introduction

Negishi cross-coupling of organohalides or pseudo-halides with Zn organometallics, commonly catalysed by Ni or Pd, has found a great many applications in synthetic chemistry, natural product synthesis and medicinal chemistry.^{1–3} The biggest competitive advantage of this cross-coupling reaction, against other popular methodologies like Suzuki–Miyaura, Corriu–Kumada or Stille couplings, lies in the reactivity of the Zn organometallic used as the transmetalation reagent. As a moderately reactive metal, that is more electronegative than Li or Mg which are often used in organometallic nucleophiles, Zn exhibits high functional group tolerance. This enables the preparation of functionalized organozinc reagents (ester, cyanide, ether, amines, *etc.*), which successfully react with a wide range of functionalized organohalide electrophiles including ketones, nitriles, nitro-compounds, esters and amides.⁴ The reduced basicity of Zn organometallics, compared to related Mg and Li species, improves their cross-coupling performance by reducing Pd deactivation under catalytic conditions, which with more basic organometallics leads to organopalladate species.^{5,6} On the other hand,

organozinc reagents are reactive enough to undergo cross-coupling reactions more rapidly than organoboron and organotin reagents, through rapid transmetalation steps, which allow for faster reactions with less-reactive substrates.^{7,8} Thus, Negishi couplings offer a good balance between selectivity, functional group tolerance and reactivity and as such are highly attractive for small-molecule synthesis.

However, the Negishi coupling presents some important challenges: the presence of unwanted reactivity, leading to organozinc homocoupling and reductive dehalogenation of the electrophilic coupling partner, alongside the need to perform reactions in rigorously dried, non-protic solvents, with protection from atmospheric moisture during the reaction, to avoid protonolysis of the organozinc reagent. This latter limitation is particularly important in industrial settings and in laboratories lacking glovebox and Schlenk equipment. In addition, the relatively high reactivity of organozinc reagents poses challenges in controlling regioselectivity when multiple sites for C–C bond formation are present on the organohalide coupling partner.

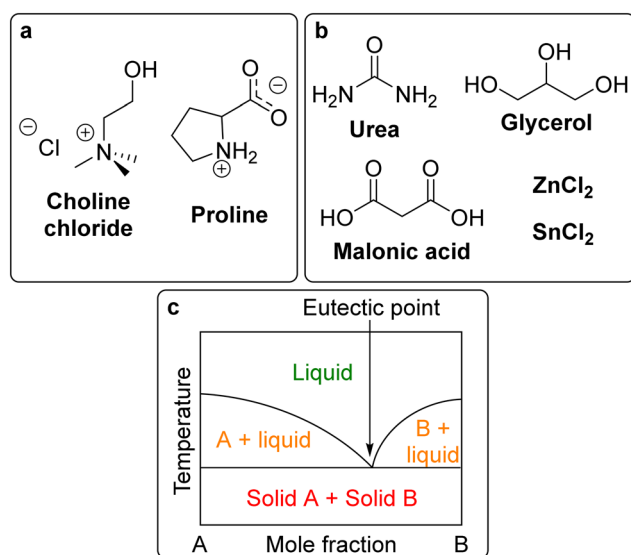
Unproductive reactivity (leading to organozinc homocoupling and reductive dehalogenation) can be partially controlled by strategies that increase the rate of reductive elimination of the product relative to the unproductive pathways. Some examples are the employment of bulky and

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electron-rich phosphines like XPhos, the use of slightly hindered electrophiles, and the use of bulky organozinc reagents.^{9,10} The requirement for dry solvents and protective atmosphere is more challenging to address, due to the inherent reactivity of organozinc reagents towards water. It is here that some surprising recent developments in alternative solvent systems (*vide infra*) prove useful.

Recent years have seen a huge amount of research aiming to develop less harmful, more sustainable chemistry. In palladium catalysis, for example, the move towards using more sustainable solvents systems has been of particular interest.¹¹ Amongst the many candidates being studied as alternatives to conventional volatile organic solvents, deep eutectic solvents (DES) have gathered significant momentum.^{12–20} DES are a diverse range of liquids formed from mixtures of Lewis basic components, or salts such as choline chloride, with Lewis acidic species or hydrogen bond donors, such as urea or glycerol.^{13,21} Eutectic mixtures of these species melt at considerably lower temperatures than their pure components (Scheme 1). They are related to ionic liquids (ILs), which have been used extensively in catalysis and as self-organised solvents for synthesis,^{22–30} since they often contain a high proportion of ions. However, they are generally cheaper and simpler to prepare, often composed of non-toxic and biodegradable components and many have the potential to be prepared from renewable resources, which makes them attractive alternatives to harmful volatile organic solvents, where there is an on-going drive for replacement. Compared to the related ILs, the simple, low *E*-factor synthetic routes to DES and their potentially reduced toxicity and environmental impact makes them particularly attractive from a green chemistry point of view. The application of DES in organic synthesis and catalysis has been reviewed recently.^{14–19,31}

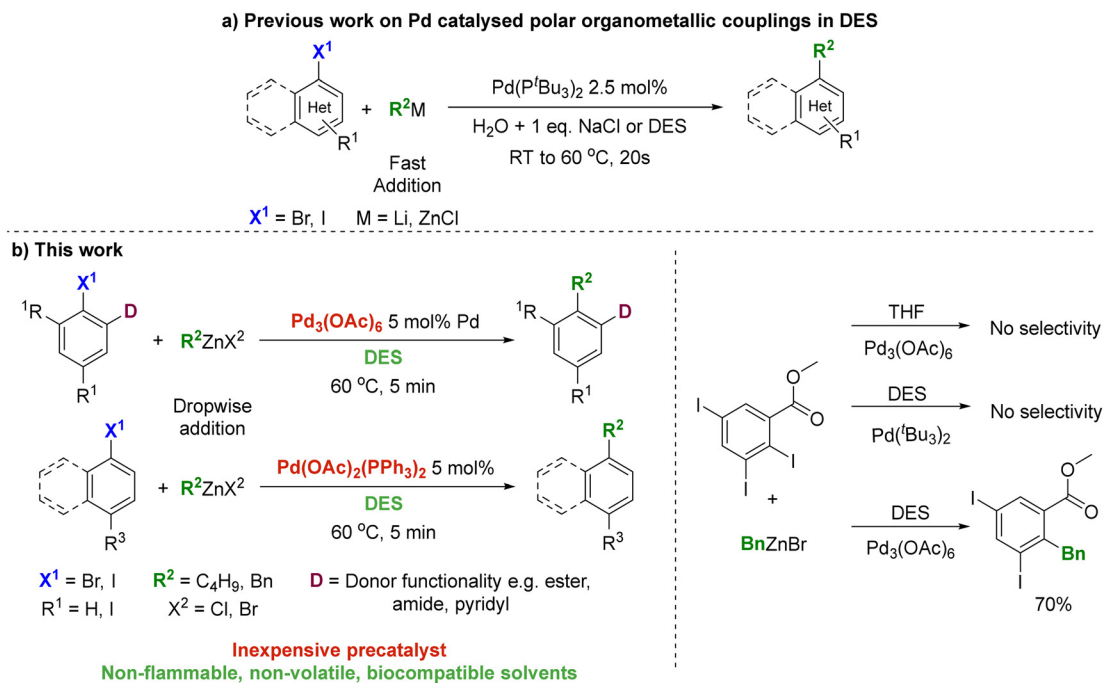


Scheme 1 Examples of (a) ionic or zwitterionic and (b) neutral components commonly found in DES. c) A schematic phase diagram for a DES.

Remarkable recent studies have shown that ammonium-based DES can be used as solvents for reactions involving highly reactive organometallic reagents, despite the presence of protic functional groups and substantial amounts of water in the DES. In foundational work it was shown that in the presence of ammonium salts and Lewis bases, the Schlenk equilibrium of Grignard reagents is shifted towards the dialkylmagnesium and magnesium halide side.³² It has also been suggested that water, as well as other Lewis bases, prefer to interact with magnesium halides rather than alkylmagnesium species. Thus, magnesium halides can exert a water-scavenging effect that provides some protection for the organometallic reagent.³³ It should be noted that in these studies, conditions somewhat similar to those in flow chemistry were used and the outcomes in batch conditions in solvents containing water or using “on water” conditions may be different. Indeed, a recent report using a range of techniques has suggested that some of these reactions may be “interface-driven”, with accumulation of reactants at the liquid surface.³⁴ Another recent report has explored the addition of Li organometallics to ketones under “on water” conditions.³⁵ These observations were followed shortly after by reports that electrophilic addition reactions involving organolithium and organomagnesium reagents could take place successfully in wet cholinium-based DESs in open air.³⁶ It was also shown that coupling of organolithium reagents with aryl halides, catalysed by Pd complexes, can proceed in DES, although solvent-free conditions were even more effective in this study.³⁷ Since then, more synthetic strategies involving air-sensitive organometallic reagents in DESs have been reported.^{38–46} The increased stability of polar-organometallic reagents in DESs, compared to volatile organic solvents, has been attributed to the participation of water molecules within the complex hydrogen-bonded network present in DES. This is thought to “deactivate” water, leading to a similar scavenging effect to the magnesium halides described above, slowing down the rate of hydrolysis of the organometallic. Indeed, some evidence for reduced water–water interactions, and increased water–DES interactions, have been seen in simulations of DES/water mixtures.⁴⁷ It has also been proposed that the halide ions found in many DES can coordinate to the organometallic reagent to form anionic “ate”-complexes with enhanced nucleophilicity, while at the same time, the hydrogen-bond donor components of the DES can bind to and activate the electrophile, increasing the rate of the desired reaction.⁴⁸

Using DES as solvents for cross-coupling reactions involving air-sensitive organometallic reagents therefore holds promise to circumvent the requirement for strictly air-sensitive conditions for Negishi couplings, allowing their use in a wider range of settings. DESs also present advantages in terms of the sustainability of the solvents. Recently, Capriati and co-workers reported cross-couplings between C(sp³) and C(sp²) lithium and zinc organometallic reagents and haloarenes in DES, or “on water” with NaCl as an additive, using Pd(P^tBu₃)₂ as the catalyst (Scheme 2a).^{49,50} Cross





Scheme 2 Summary of previous methodologies, and that used in the present study, for Pd-catalysed cross-couplings with reactive polar organometallic reagents in DES. R^3 is a range of substituents.

coupling with the organolithium reagents did not proceed in DES, where only reductive dehalogenation of the substrate was observed. However, “on water” the reactions were successful and tolerant of haloarenes functionalized with both electron-withdrawing and donating functional groups. Organozinc reagents underwent cross-coupling satisfactorily in both DES and “on water”, but DES were more suitable for substrates bearing hydrogen-bond donors *e.g.*, amines, carboxylic acids, and alcohols, since these haloarenes did not undergo coupling on water. The authors found that the reactions “on water” were very sensitive to the nature of the pre-catalyst, temperature and organozinc stoichiometry. In related work, Negishi coupling reactions between aryl bromides and $\text{C}(\text{sp}^3)\text{-X}$ ($X = \text{Br, I}$) halides have been reported “on water” with the assistance of micelles formed by non-ionic surfactants.^{51–54} Remarkably, in these systems it was possible to form the organozinc reagent *in situ* in the hydrophobic environment of the micelles, rather than the organometallic being formed *ex situ* and added at the start of the reaction. The synthesis of the Zn organometallic exhibited good functional group tolerance and the cross-coupling reaction was successful for aryl halides functionalized with ketones, esters and methoxy groups. Nitriles and alcohols were not tolerated. Since this publication, some additional cross-coupling reactions have been reported in DES, including Mizoroki–Heck from Dilauro *et al.* to give functionalised THF derivatives in good yield.⁵⁵ D’Amico *et al.* reported Pd-catalysed cross-coupling of aryl bromides with functionalised thiophene derivatives in choline chloride/glycerol DES.⁵⁶ Antenucci *et al.* reported Pd-catalysed coupling of arenediazonium *o*-benzenedisulfonimides and

organolithium reagents in DES to give biaryl products.⁵⁷ Niakan *et al.* have also reported Pd-free Suzuki and Sonogashira cross-coupling in DES, with immobilised Co nanoparticles acting as the catalyst.⁵⁸ Most recently, Ramos-Martin *et al.* reported Friedel–Crafts benzylations in DES, catalysed by FeCl_3 .⁵⁹

Although DES have been shown to be promising alternative solvents for Negishi coupling,⁵⁰ the role of the Pd pre-catalyst has not yet been reported. This has relevance for the potential application of these methodologies, where the use of cheaper, air-stable and readily available Pd pre-catalysts, like $\text{Pd}_3(\text{OAc})_6$, would maximise their utility in a range of settings. An understanding of pre-catalyst activity towards different substrates, and selectivity in the presence of multiple carbon-halide or pseudohalide bonds in the substrate, is also important. This contribution will present methodologies that address both of these challenges, in addition to allowing a regioselective Negishi coupling in DES for the first time, whilst maintaining the ability to perform the cross-coupling reactions in wet solvents with no protective atmosphere (Scheme 2b).

Results and discussion

Developing Negishi coupling in DES using $\text{Pd}_3(\text{OAc})_6$ as the pre-catalyst

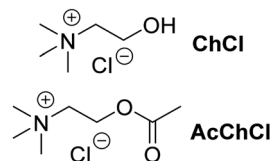
Previous work demonstrated that the Pd(0) pre-catalyst $\text{Pd}(\text{P}^t\text{Bu}_3)_2$ allows highly effective Negishi coupling of a range of substrates in DES.⁵⁰ However, this complex is air-, moisture- and temperature-sensitive and is ideally handled under an inert atmosphere and stored in a glovebox at low



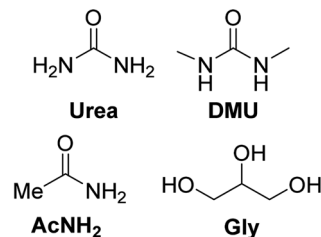
temperature. This negates some of the benefits of being able to perform these reactions in DES on the open bench, without protective conditions. Our aim was to develop a protocol that allows related reactions to take place using cheaper, air- and moisture-stable Pd(II) precursors based on ubiquitous Pd₃(OAc)₆. To develop appropriate reaction conditions we studied the reaction of ⁿBuZnCl with methyl 2-iodobenzoate under similar conditions to those reported by Marder and Lei (Table 1).⁶⁰ This reaction was chosen for several reasons: it was reported to proceed rapidly (less than 2 min for full conversion) and consideration of recent work with DES suggested that fast organometallic reactions were more likely to be compatible with the water and functional groups present in DES. Additionally, no phosphine ligand is required, which reduces the potential for ligand solubility problems in the DES and the high costs associated with some phosphines.⁶¹ In general, fewer additives is desirable in terms of sustainability, cost and convenience. Finally, the reaction is compatible with ammonium halides, which were used in the formation of the DES.

To explore solvent effects, five DES were prepared as described in the SI and shown in Scheme 3. Four were based on the hydrogen-bond acceptor (HBA) choline chloride (ChCl) with glycerol (Gly), urea, acetamide (AcNH₂) and dimethylurea (DMU) as the hydrogen-bond donor (HBD) components (1:2 ratio of HBA:HBD). One was based on acetylcholine chloride (AcChCl) as the HBA, with urea as the

a) Investigated salts



b) Investigated hydrogen-bond donors



Scheme 3 Salts, hydrogen-bond donors, and their abbreviations, used to prepare the DES studied here.

HBD (1:1 ratio). ChCl-based DES, with urea and glycerol, were chosen since these are the most widely used DES solvents for organic transformations and they are comparatively well studied and readily available. DES based on acetamide, dimethylurea and AcChCl were devised to reduce the protic nature of the HBD fraction and modify the solubility of the reagents used in the catalytic reactions. For ChCl/urea and ChCl/Gly, batches of DES were prepared under

Table 1 Summary of initial reaction screening and optimisation using the pre-catalyst Pd₃(OAc)₆

Entry	Solvent	Solvent wet/dry	Temp. (°C)	Pd ₃ (OAc) ₆ (mol% Pd)	ⁿ BuZnCl (equiv.)	RZnCl rate of addition	Yield of 2a (%)	Yield of 3 (%)
1 ^{a,c}	THF	Dry	40	2	2	Dropwise	84	1
2 ^a	ChCl/urea	Dry	40	1	1.6	Dropwise	88	10
3 ^a	ChCl/urea	Wet	40	1	1.6	Dropwise	27	3
4 ^a	ChCl/urea	Dry	40	2	1.6	Dropwise	90	6
5 ^a	ChCl/urea	Wet	40	2	2.0	Dropwise	55	5
6 ^a	ChCl/urea	Dry	40	2	2.0	Fast, single	31	10
7 ^a	ChCl/urea	Wet	40	2	2.0	Fast, single	46	8
8 ^a	ChCl/urea	Dry	60	1	2.0	Dropwise	87	12
9 ^a	ChCl/urea	Wet	60	1	2.0	Dropwise	80	5
10 ^a	ChCl/urea	Wet	60	2	2.0	Dropwise	97	0
11 ^a	ChCl/urea	Wet	60	2	2.0	Fast, single	70	7
12 ^a	ChCl/urea	Dry	60	0	2	Dropwise	0	0
13 ^{a,d}	ChCl/urea	Dry	60	2	0	—	0	0
14 ^b	ChCl/DMU	Wet	80	1.5	2.0	Dropwise	14	1
15 ^b	ChCl/AcNH	Wet	80	1.5	2.0	Dropwise	14	1
16 ^a	ChCl/G	Wet	40	0.5	2.0	Dropwise	6	0
17 ^{a,d,e}	ChCl/G	Wet	40	0.5	2.0	Dropwise	0	0
18 ^{a,d}	ChCl/G	Dry	40	1	1.6	Dropwise	0	0

Typical conditions: **1** (1.0 mmol) and Pd₃(OAc)₆ were allowed to mix for 10 minutes in 5 cm³ solvent at the desired temperature. **a** in 2 cm³ dry THF was added (either dropwise or in one quick addition) while the reaction mixture was stirred at 1000 rpm, then the reaction was stirred for 5 minutes prior to workup (see SI for details). ^a NMR yields determined by ¹H NMR spectroscopy. ^b GC yields determined by gas chromatography. ^c 1 equiv. [Bu₄N]Br added as additive. ^d Reaction period was 12 h. ^e 10 cm³ solvent used.



conditions designed to minimize their water content and these were stored under N₂. These DES are denoted “dry” and typical water contents for these materials were around 400 ppm. In addition, batches of each DES were prepared directly from the components as received and stored in air. These DES are denoted “wet” and typical water contents for these were between 1 and 5 wt%, depending on the HBD used. Water contents for each DES, determined by Karl Fischer titrations, are given in the SI. Selected results from an initial screening study for the conversion of **1** into **2a** and **3** as a function of solvent and reaction condition are shown in Table 1. A full set of data are given in the Table S1.

In THF, under an inert atmosphere of N₂, the reaction proceeded smoothly (Table 1, entry 1). Formation of **2a** is accompanied by dehalogenation of **1** to form **3**, through a competing β-hydrogen elimination pathway. Although this is a relatively minor side reaction in THF, it can become more significant under some conditions in DES (*vide infra*). When comparing the results from different DES, it was found that those based on glycerol were not suitable reaction media, with yields of **2a** of 6% at most and unreacted **1** recovered (Table 1, entries 16–18). **3** was not detected at the end of the reaction for either the “wet” or “dry” ChCl/Gly DES. In all cases, a white precipitate was formed during the reaction and there was no visible evidence of black Pd(0) particle formation, which occurs in THF. Although these DES have been shown to be compatible with other organometallic reactions,⁶² and when using Pd(P^tBu₃)₂ as pre-catalyst for Negishi couplings,⁵⁰ in this case it appears that the alkylzinc reagent degrades more rapidly in the polyol-based DES than it takes part in the coupling reaction with Pd₃(OAc)₆ as pre-catalyst. Reasoning that this reactivity was likely to be promoted by the OH groups on the HBD in ChCl/Gly, which are both reactive and strong H-bond donors, the reaction was then studied in DES based on urea and related HBDs. These DES were found to support productive catalysis under a range of conditions.

Because of the rapid degradation of RZnX reagents by water, one might imagine that using “dry” DES would inherently be more successful than using the “wet” solvents. However, our results show a more complex picture for this reaction. At 40 °C, with dropwise addition of the ⁿBuZnCl, a significantly improved yield of **2a** is seen for the dry DES compared to the wet DES (Table 1, entries 2 to 5). However, when the organozinc reagent is added in a single, rapid addition the yield of **2a** is similar between the dry and wet DES, with a somewhat higher yield for the wet solvent, although both show relatively poor yields under these conditions (Table 1, entries 6 and 7). Dilauro *et al.* have previously shown that a rapid, single addition of an organometallic reagent to reactions taking place “on water” or in DES promotes catalysis,^{49,50} but this seems less effective in our system, with dropwise addition generally producing better results.

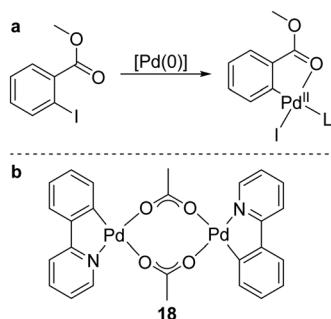
By increasing the reaction temperature from 40 to 60 °C it was possible to circumvent the need for dry DES and

protective N₂ atmosphere (Table 1, entries 8 and 9). Further optimisation found that a slightly higher catalyst loading (2 mol% Pd) also increased the yield at both 40 and 60 °C in the wet solvent. As such, Table 1, entry 10 represents optimised conditions for this reaction, which allowed highly effective catalysis (up to 97% yield of **2a**) in “wet” ChCl/urea (typically around 3 wt% water), without any precautions needed for the preparation or storage of the DES or any protection of the reaction from atmospheric water or O₂. The reaction was repeated a number of times to give an average yield of **2a** of 86% over 6 runs (Table S2, entries 14.1–14.6), demonstrating good reproducibility. A double-scale reaction gave an isolated yield of **2a**, after workup and purification, of 80%, giving confidence that NMR/GC yields were representative of the isolated yields achievable following purification. Investigating the rate of addition of **a** under the optimised conditions, a fast, single addition of the organozinc reagent led to a degradation in performance (Table 1, entry 11). Control experiments under these optimised conditions, without addition of Pd, and without organozinc reagent (Table 1, entries 12 and 13) led to no formation of **2a**, and no formation of dehalogenated side product **3**, confirming that catalytic Pd is required and that in the absence of **a**, there is no dehalogenation of the substrate.

Exploration of additional DES as reaction media generally led to poor yields of **2a** (Table 1, entries 14 and 15). This is interesting when the solubility of the reagents in the DES is considered, as there seems to be no simple correlation between the solubility of the reactant **1** and the success of the reaction. Pd₃(OAc)₆ has good solubility in all of the DES. **1** has only partial solubility in ChCl/urea, but the reaction proceeds very effectively under the optimised reaction conditions regardless. Moving to ChCl/AcNH₂ and ChCl/DMU improves the solubility of the aryl iodide, but the yield of **2a** is substantially reduced and evidence of reactions between **1** and the HBD are observed (Table S1). In AcChCl/urea the solubility of **1** and Pd₃(OAc)₆ is very poor, and so no catalytic experiments were carried out in this DES. When the above observations are coupled with the poor yields seen in DES based on Gly as the HBD, it is clear that the factors influencing the choice of DES for reactions such as this are complex and some trial-and-error testing is currently necessary to determine the optimal DES for the reaction, even amongst chemically similar types.

Finally, re-use of the DES/Pd mixture was tested by reaction of **1** with **a** to give **2a** (conditions as per Table 1, entry 10), followed by extracting the organic products and unreacted starting material directly from the DES using dichloromethane. The results showed a decrease in the yield of **2a** from 83% for the first cycle, to 64% for the second and 56% after a third cycle (see SI for details). There is clearly degradation in performance across these runs and so catalyst/solvent re-use is not attractive for this system.





Scheme 5 a) Proposed *ortho*-coordination between methyl 2-iodobenzoate and palladium following oxidative addition, and prior to any isomerisation. Pd speciation in the DES is likely to be complex. There are many species capable of coordinating at position L (e.g. $[\text{OAc}]^-$, Cl^- , I^- , urea, H_2O , $[\text{Ch}]^+$). b) Palladacycle **18** prepared to explore metallocycle reactivity.

We tentatively propose that this type of stabilisation may facilitate pre-coordination of the substrate to Pd(0) and/or a faster oxidative addition in this system, which allows the rate of productive catalysis to outcompete the rate of hydrolysis of the organozinc reagent in the DES. This is at odds with previously measured kinetics for the oxidative addition of related *ortho*-donor functionalised aryl iodides to Pd(0), in the presence of phosphine ligands, where *ortho*-donors actually reduced the rate of oxidative addition.⁶⁴ This was attributed primarily to steric effects. However, the presence of preferential Pd-phosphine interactions in the previous work may influence the observations and in the present study, where no phosphines are present, the *ortho*-donor interactions may be sufficiently stabilising to either allow pre-coordination of the substrate to Pd(0) or to otherwise facilitate oxidative addition. Interestingly, when iodo-2-phenylpyridine is the substrate, the yields are reduced for reactions with $^n\text{BuZnCl}$ as the nucleophile, while good yields are maintained with BnZnBr , despite the potential for formation of a related 5-membered palladacycle in both cases (**18**, Scheme 5b). In this case, the larger pyridyl substituent may begin to inhibit oxidative addition for steric reasons, as seen in the Pd(0) phosphine systems studied by Jutand.⁶⁴ A subtle balance of competing effects is likely to be at play here.

To explore the potential role of palladacyclic intermediates, (**18**, Scheme 5b), a known and stable metallocycle was synthesised *ex situ* as a model for an intermediate of the reaction that yields product **10b**.⁶⁵ This was reacted with a stoichiometric amount of benzyl zinc bromide under similar conditions to the optimised catalytic procedure above (see SI for details). This led to formation of product **10b**, albeit with only a modest yield of 17%, with a significant amount of the alkylzinc halide homocoupling product, bibenzyl, produced as a side product under these conditions. **18** was also used as a pre-catalyst for the coupling of **1** with BnZnBr , giving an overall yield of 89% **2a** (see SI for details), which is comparable to when $\text{Pd}_3(\text{OAc})_6$ is used as the pre-catalyst. These experiments demonstrate that palladacycles, such as **18**, can convert to product under

catalytically relevant conditions and are competent pre-catalysts, supporting their potential role as catalytic intermediates.

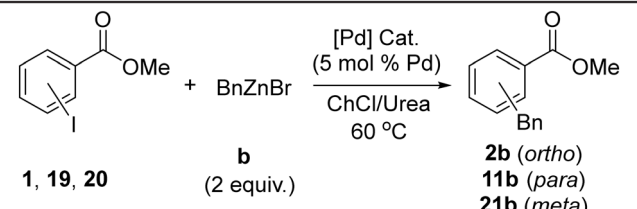
Reactions between 2-iodoanisole and **a** or **b** showed poor yields, despite the potential formation of a 4-membered palladacycle *via* coordination of OMe to Pd. This result suggests that these phosphine-free Negishi reaction conditions in ChCl /urea are not very effective for electron-rich iodoarenes, which are expected to undergo oxidative addition to Pd more slowly. Activity was also found with bromoarenes, albeit with only modest yields. Thus, when methyl 2-bromobenzoate reacted with **b**, the yield of product **2b** was 44%. It was initially thought that bromoarenes would undergo oxidative addition to Pd too slowly and so productive catalysis would be unable to compete with the rate of hydrolysis of the organozinc reagent in the DES. However, it seems that this effect is not strong enough to completely suppress the catalytic reaction.

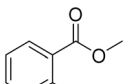
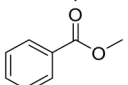
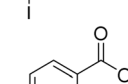
Addition of PPh_3 to the system, by pre-formation of the Pd complex $\text{Pd}(\text{OAc})_2(\text{PPh}_3)_2$ and its use as pre-catalyst, allowed for good yields of *para*-substituted iodoarenes in most cases (Scheme 4, green box). However, reactions involving iodoarenes functionalised at the *ortho* position generally exhibited poor yields under these conditions, so this represents a complementary methodology to that using $\text{Pd}_3(\text{OAc})_6$, as described above. With $\text{Pd}(\text{OAc})_2(\text{PPh}_3)_2$, the differences between electron-rich and electron-poor iodoarenes were no longer appreciable. Jutand and co-workers showed that this Pd(II) pre-catalyst is reduced *in situ* by an intermolecular cleavage of a phosphine and one of the acetate groups to yield a very reactive and unstable low-valent Pd species.⁶⁶ Recently, Fairlamb and co-workers characterised this as a Pd(I) dimer bridged by one acetate and one phosphine ligand, which forms a very active trinuclear Pd species upon oxidative addition with haloarenes.⁶⁷ A similar Pd(I) dimer, bridged by bromide, was prepared by Colacot *et al.* by reduction of $\text{Pd}(\text{COD})\text{Br}_2$ in the presence of 1 equivalent of P^tBu_3 and a base.⁶⁸ The mechanism for the reduction of Pd(II) in this system was later explored in detail.^{69,70} Although Pd speciation in the DES is currently unknown, similar activation processes and formation of highly active Pd species derived from the pre-catalyst is possible and is consistent with the observed increase in reactivity towards *para*-substituted iodoarenes. However, in a chloride-rich environment, such as that provided by the DES, the formation of anionic Pd “ate” complexes is also possible.^{71–73} In addition, it has been proposed by Jutand and co-workers, that in chloride-rich environments, chloride can bind to Pd and narrow the P–Pd–P angle in bisphosphine complexes, which may also enhance rates of oxidative addition of iodoarenes to Pd.^{71,74}

The effect of pre-catalyst on selectivity

The complementarity in the two methodologies shown in Scheme 4, where different pre-catalysts allow the use of aryl



Table 2 Summary of results of cross-coupling reactions with isomers of methyl iodobenzoate and benzylzinc bromide using different pre-catalysts and conditions


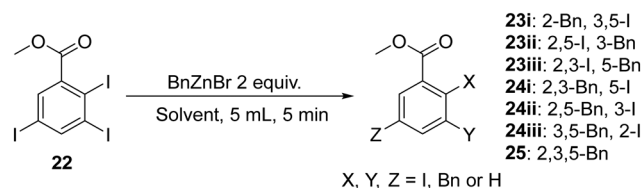
Substrate	Pd ₃ (OAc) ₆ (% yield)	Pd(OAc) ₂ (PPh ₃) ₂ (% yield)	Pd(P ^t Bu ₃) ₂ (% yield)	
			Dilauro <i>et al.</i> conditions ^{a,b50}	Conditions from this work
	84	61	98	63
	43	52	94	93
	60	71 ^b	75	82

^a [Pd] cat. at 2.5 mol% Pd. ^b BnZnBr at 3 equiv., as shown in Scheme 4. GC yields determined by gas chromatography.

iodide substrates with different substitution patterns, inspired us to investigate the effect of Pd pre-catalyst towards *ortho*-, *meta*- and *para*-substituted methyl iodobenzoates **1**, **19** (*para*) and **20** (*meta*) in more detail. Table 2 shows a summary of the % yield of the desired C–C coupled product **2b**, **11b**, and **21b** (*meta*), from isomers of methyl iodobenzoate with BnZnBr in ChCl/urea using different pre-catalysts and reaction conditions.

The results in Table 2 show significant variation in the amount of C–C coupled product that is formed depending on the substitution pattern of the substrate, the pre-catalyst and the conditions used. When Pd₃(OAc)₆ is used, higher yields for *ortho*-, rather than *meta*- or *para*-iodoarene, as described above, are clear. Pd(OAc)₂(PPh₃)₂ is less active for these substrates, and more consistent in yield for the different substitution patterns, albeit with a preference for *para*-substituted substrates. Pd(P^tBu₃)₂ is the most active catalyst for these reactions in ChCl/urea, but its activity depends on the reaction conditions used. Under the conditions reported by Dilauro *et al.*,⁵⁰ good to excellent yields are seen for the *para*- and *ortho/meta*-substituted substrates respectively. If the optimised reaction conditions from the present study (Scheme 4) are used, which involve dropwise addition of the alkylzinc reagent, rather than a fast single addition, and a slightly higher catalyst loading, Pd(P^tBu₃)₂ performs less effectively for the *ortho*-substituted iodoarene, but similarly for the others. This is the opposite of what is seen for Pd₃(OAc)₆ and highlights the subtle effects at play in these reactions.

As different pre-catalysts show differing activities towards aryl iodide substrates with different substitution patterns, this inspired us to investigate a regioselective Negishi coupling in DES where multiple C–X bonds are present in the electrophile. Control over the site of C–C bond formation in such substrates is highly desirable, allowing sequential functionalisation at different positions without the need for protection/activation, thus simplifying construction of complex molecular architectures. To this end we aimed to leverage the selectivity for C–C bond formation *ortho* to a coordinating functional group (*e.g.* ester, amide) when Pd₃(OAc)₆ is used as a pre-catalyst. Carboxylate-directed Suzuki–Miyaura coupling has been reported previously,⁷⁵ and Pd(I) pre-catalysts have been shown to allow chemoselective Kumada and Negishi cross coupling for substrates containing two or more different halides/pseudohalides,⁷⁶ but to the best of our knowledge donor-group-directed, regioselective Negishi reactions have not been described.



Scheme 6 Regioselective Negishi coupling in ChCl/urea. Reactions performed at 60 °C in ChCl/urea, and at 40 °C in THF, with 1 equiv. of [NBu₄]Br, as per ref. 60, as a control reaction. Selected products, not including dehalogenated species, and their compound numbers are shown.



Methyl 2,3,5-triiodobenzoate (**22**), with one C–I bond *ortho* and two *meta* to the ester, was coupled with BnZnBr, under our optimised reaction conditions for *ortho*-coordinating functional groups in ChCl/urea (Scheme 6). More than twenty potential products, including regioisomers, are possible in this reaction *via* substitution of one or more I for Bn, and/or *via* dehalogenation. As such, this substrate represents a challenging test for selectivity in this reaction. Under these conditions with BnZnBr, 88% of the electrophile was consumed, resulting in 70% yield of the mono-substituted product **23i**, where Bn has added to the position *ortho* to the ester. Only 2% yield was observed for the other mono-substituted isomers. Small amounts of di-substituted **24i–iii** and tri-substituted **25** cross-coupled products were observed at 4% and 5% yield respectively. The mono-dehalogenated side products were present at 1%. Finally, a combination of partially substituted products with different degrees of dehalogenation were produced in 6% yield (Scheme S4, Table S6). Overall, the 70% selectivity of the reaction towards one, out of many potential products was very good. In order to confirm the connectivity of the major, monosubstituted product, **23i** was isolated by column chromatography and analysed by NMR spectroscopy (Fig. S11–S16). A chemical shift of δ 8.37 ppm for the proton at the *para* position of the aromatic ring suggested that it is surrounded by two iodine atoms. The only connectivity pattern which could account for such downfield displacement is the one in which the new C–C bond is formed *ortho* to the ester. In addition, in the HMBC spectrum, a distinctive four-bond correlation between the methylene protons on the Bn group and the carbonyl carbon atom was detected. Both features are consistent with the proposed formulation of **23i**.

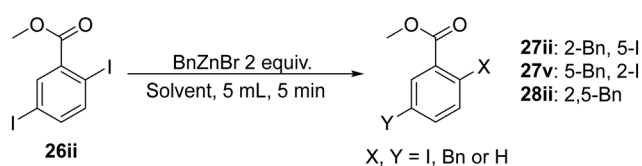
To explore the role of the DES in this regioselectivity, a control reaction was performed in THF using reaction conditions similar to those originally reported by Lei.⁶⁰ In dry THF in the presence of 1 equiv. [NBu₄]Br and 5 mol% Pd₃(OAc)₆, **22** was reacted with BnZnBr at 40 °C. In contrast to the DES, the reaction in THF resulted in a yield of **23i** of only 10%, in addition to 28% of the other mono-substituted isomers. The di-substituted products were formed in 35% yield, while the tri-substituted product was present at 15%. Clearly the regioselectivity is poor in THF. Surprisingly, the degree of dehalogenation was also higher for this substrate in THF than in the DES, with a total of 13% of a range of dehalogenated products being formed.

The impact of pre-catalyst on the regioselectivity of this reaction in ChCl/urea was also explored. Using Pd(P^tBu₃)₂, under the conditions optimised by Dilauro *et al.*,⁵⁰ 84% of the electrophile **22** was consumed, now resulting in 45% yield of *ortho*-substituted product **23i**, with 2% yield of monosubstituted **23ii–iii**. The tri-substituted **25** cross-coupled product was observed at 7% yield and di-substituted products **24i–iii** at 11%, alongside some bis-dehalogenated (**1**, **20**) side products at 3% (Scheme S4, Table S9). Using Pd(P^tBu₃)₂ as pre-catalyst, but under the

reaction conditions used for the Pd₃(OAc)₆ reaction above, led to even lower electrophile consumption and less selectivity (Scheme S4, Table S8). Finally, the reaction between **22** and BnZnBr was repeated with Pd(OAc)₂(PPh₃)₂ as pre-catalyst, leading to 60% of the electrophile being consumed and 44% yield of **23i**. Some formation of the tri-substituted product **25** was observed at 3% yield, alongside the mono-dehalogenated (**26i–iii**) and bis-dehalogenated (**1**, **20**) side products at 5% and 7% respectively (Scheme S4, Table S7).

From this series of reactions, it can be seen that the reaction catalysed by Pd₃(OAc)₆ in DES shows the highest yield of, and best selectivity for, the *ortho*-monosubstituted product **23i**. When PPh₃ is present, with pre-catalyst Pd(OAc)₂(PPh₃)₂, the yield is lower at a similar reaction time point, although some selectivity towards **23i** is retained. Although this pre-catalyst was introduced to allow coupling of *para*-substituted iodoarenes under our conditions (Scheme 6), it proved to also be moderately competent for some of the listed *ortho*-substituted substrates. Pd(P^tBu₃)₂ is a more active pre-catalyst for Negishi coupling reactions in DES, including those of *meta*-substituted haloarenes,⁵⁰ but this activity now leads to much lower selectivity with electrophile **22**, as a higher proportion of di- and tri-substituted products form than with the other pre-catalysts. Moving from the DES into THF leads to a loss of all selectivity in the reaction.

To further investigate selectivity, diiodoarene **26ii** was synthesised, and the selectivity conferred by the three pre-catalysts investigated (Scheme 7). When Pd₃(OAc)₆ was used, 83% of the electrophile was consumed, with 64% yield of the *ortho*-substituted product **27ii** and 16% yield of the disubstituted product **28ii** alongside some dehalogenated products (Scheme S5, Table S10). With Pd(OAc)₂(PPh₃)₂ as the pre-catalyst, 43% of the electrophile was consumed to give 20% of the *ortho*-substituted product **27ii** and 5% yield of the *meta*-substituted product **27v**. Disubstituted product **28ii** was now present at 4% alongside some dehalogenated products (Scheme S5, Table S11). When the reaction was carried out with Pd(P^tBu₃)₂ under the conditions referred to earlier in this work, 72% of the electrophile was consumed to give 16% *ortho*-substituted product **27ii** and 47% disubstituted product **28ii**, in addition to some partial dehalogenation products (Scheme S5, Table S12). When carried out under the conditions optimised by Dilauro *et al.* though, 84% of the electrophile was consumed to give 25%



Scheme 7 Regioselective Negishi coupling in ChCl/urea. Reactions performed at 60 °C in ChCl/urea. Selected products, not including dehalogenated species, and their compound numbers are shown.



ortho-substituted product **27ii** and 44% disubstituted product **28ii**, alongside some dehalogenated products (Scheme S5, Table S13).

Overall, these reactions with diiodoarene **26ii** follow similar trends to those seen for triiodoarene **22**, with Pd₃(OAc)₆ giving the highest selectivity for a single substitution, at the *ortho*-position. Pd(OAc)₂(PPh₃)₂ led to very little yield, although no disubstitution was observed. When Pd(P^tBu₃)₂ was used, a preference for disubstitution at both the *ortho*- and *meta*-positions was seen, which is consistent with its higher activity, which delivers lower selectivity as seen before. The results for both systems are summarised in Table 3.

Summary and conclusions

Novel methodologies for Negishi cross coupling of aryl iodides with alkylzinc halides in DES based on simple pre-catalysts, which are not temperature-, air- or moisture sensitive, have been developed. Two complementary methodologies, either phosphine-free with Pd₃(OAc)₆ as the pre-catalyst, or where PPh₃ is added as a supporting ligand, allow cross coupling to take place with a range of aryl iodides. Substrates with donor groups *ortho* to the C-I bond, e.g. ester, amide, pyridyl, which can coordinate to metals present in the system, do not require a phosphine ligand to be added to achieve good yields. Palladacycle formation as part of the catalytic mechanism is proposed in these un-supported reactions, with both stoichiometric and catalytic testing of an isolated palladacycle supporting this. To functionalise substrates *meta* and *para* to the C-I bond requires the addition of PPh₃ as a supporting ligand.

The effect of Pd pre-catalyst on reactivity and selectivity towards iodoarenes with different substitution patterns

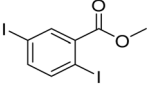
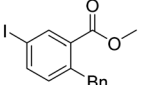
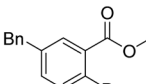
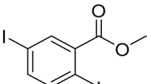
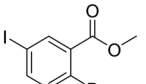
was explored, using Pd₃(OAc)₆, Pd₃(OAc)₂(PPh₃)₂ and Pd(P^tBu₃)₂ pre-catalysts. Pd(P^tBu₃)₂ is the most active of these, and as demonstrated previously,⁵⁰ achieves the widest substrate scope, but is the least selective when multiple cross-coupling sites exist on the iodoarene. Pd₃(OAc)₆ is active towards *ortho*-donor substituted substrates and allows highly selective cross-coupling when more than one C-I bond is present in the substrate. Pd₃(OAc)₂(PPh₃)₂ is the least active catalyst for *ortho*- and *meta*-substituted substrates, but competent for *para*-substitution, which allows some selectivity, albeit with relatively low yields in some cases.

The complementarity of these methodologies allowed us to develop regioselective Negishi coupling in ChCl using Pd₃(OAc)₆ as pre-catalyst, where C-C bond formation *ortho* to a coordinating group was strongly preferred, even when other C-I bonds were available on the aryl halide for functionalisation. A comparable reaction in THF led to a mixture of products with no selectivity. This is an unusual example of solvent-induced selectivity, whose origin appears to lie in the relative rates of productive catalysis vs. the rate of hydrolysis of the organozinc reagent in the DES. With Pd₃(OAc)₆ as pre-catalyst, the rate of C-C bond formation *ortho* to a coordinating group, such as an ester, is fast enough to out-compete hydrolysis of the organozinc reagent. However, cross coupling at the *meta*- and *para*-positions is slower than organozinc hydrolysis in the DES and these pathways are less favoured. Overall, here and in other recent studies, DES have been shown to be intriguing alternative solvents for organometallic chemistry.

Experimental

Full experimental details are given in the SI.

Table 3 Summary of cross-coupling yields in reactions with polysubstituted aryl iodides using different pre-catalysts and conditions. Selected products are shown, see the SI for full product distributions

Substrate	Product	Pd ₃ (OAc) ₆	Pd(OAc) ₂ (PPh ₃) ₂	Pd(P ^t Bu ₃) ₂	
				Dilauro <i>et al.</i> conditions ^{a50}	Conditions from this work
		64	20	25	16
		16	4	44	47
	Starting material	17	57	16	28
		70	45	45	20
	Polybenzylated products	11	3	18	13
	Starting material	12	40	16	38

[Pd] cat. at 5 mol%, BnZnBr at 2 equiv., dropwise addition. ^a [Pd] cat. at 2.5 mol% Pd. BnZnBr at 3 equiv., as shown in Scheme 4. GC yields determined by gas chromatography.



Author contributions

L. Mauricio Murillo-Herrera: investigation, methodology, writing – original draft, writing – review & editing, funding acquisition. Spencer Frederick: investigation, methodology, writing – original draft, writing – review & editing. Scott Hicks: methodology, writing – review & editing. Simon Duckett: supervision, writing – review & editing, funding acquisition. John Slattery: conceptualization, supervision, writing – original draft, writing – review & editing, funding acquisition.

Conflicts of interest

There are no conflicts to declare.

Data availability

The data supporting this article have been included as part of the supplementary information (SI).

Supplementary information: experimental details. General methods. Preparation of all reagents. Procedures for all experiments. Characterisation data, including NMR spectra. GC methodology. Gas chromatograms. Additional raw data supporting this research is openly available from the research data repository of the University of York at <https://doi.org/10.15124/a033ed32-ff01-4c85-a1de-8955d48b0891>. See DOI: <https://doi.org/10.1039/d5cy01526j>.

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