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Support effect on Ni-based mono- and bimetallic catalysts in CO₂ hydrogenation†

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Aiming at a comprehensive understanding of support effects on Ni-based bimetallic catalyst for CO₂ hydrogenation, spectroscopy (DRIFTS) with CO as a probe molecule and temperature-programmed techniques were used to investigate the impact of different supports (MgO, CeO₂, ZrO₂) on Ni- and Ni,Fe catalysts. Kinetic parameters revealed that the higher selectivity to methanation for Ni and Ni,Fe supported on the reducible oxides (CeO₂, ZrO₂) is due to the inhibition of reverse water-gas shift reaction (RWGS) by hydrogen. A promoting effect of Fe on Ni was only observed on MgO-supported catalysts. In situ DRIFTS with CO adsorption showed different electronic properties of Ni sites with partially reduced oxide (*i.e.* ZrO₂ and CeO₂). H₂-TPR and CO₂-TPD confirmed the significant role of metal-support interaction (MSI) in CeO₂-supported catalysts for CO₂ activation. The MSI between Ni/Ni,Fe and reducible supports are crucial for catalytic performance, ultimately leading to the higher activity and stability in CO₂ hydrogenation.

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1. Introduction

CO₂ hydrogenation is a cornerstone of sustainable “power-to-gas (P2G)” concepts^{1–5} and supported Ni catalysts are considered as the most promising base metal catalyst towards methanation due to their outstanding catalytic performance and low cost.^{6–10} Despite the long history of research on nickel-based catalysts, there is still a need for a better fundamental understanding and for further rational optimization of such methanation catalysts.^{11,12}

On the one hand, regarding catalyst composition variation or promotion, partial substitution of Ni by a second metal was shown to improve the catalytic performance for CO₂ methanation.^{13,14} Compared to Ni catalysts, Ni,Fe bimetallic catalysts reach a significantly improved catalytic activity.^{15–19} For example, Grunwaldt and co-workers reported that a small addition of Fe to Ni/Al₂O₃ catalysts will increase activity and stability with an optimum ratio of Ni : Fe = 3 : 1.^{20–22} They proposed that the redox cycle of iron species on Ni nanoparticles will promote the dissociation of CO₂ and further improve the activity.²³ Similarly, Huynh *et al.* reported a Ni : Fe ratio of 4 : 1 to be optimal.²⁴ It was also reported that CO₂ is hydrogenated

to HCOO* intermediates with a lower energy barrier. Meanwhile, other reports claim that the Ni-FeO_x interfaces are more active for RWGS reaction due to its weak binding to CO* intermediates.²⁵

On the other hand, CO₂ hydrogenation over supported Ni catalysts is highly structure-sensitive.^{26–28} The size effect of Ni nanoparticles on CO adsorption has been proven to steer the selectivity of CO₂ hydrogenation, either to methane or to CO.^{29,30} In addition, the support also plays as an important role for Ni catalysts in CO₂ methanation,^{31–36} including its impact on the size and shape of the Ni nanoparticles (*i.e.* dispersion and morphology),³⁷ but also on CO₂ activation by the basicity and reducibility of the support.^{38–40} Oxygen vacancies can be formed in reducible oxides during CO₂ hydrogenation, affecting the catalyst behavior. For example, in a series of Ni catalysts on different supports (Al₂O₃, ZrO₂, CeO₂), Ni supported on non-reducible Al₂O₃ shown the lowest CO₂ conversion.³¹ Furthermore, strong metal-support interaction (SMSI) can affect the stability of CO₂ methanation, but also the selectivity.⁴¹ Such effects typically are dynamic and depend on the reaction conditions and catalysts pre-treatments in a complicated manner. Monai *et al.* recently reported that such SMSI can generate interfacial sites on Ni/TiO₂ catalysts favoring carbon-carbon coupling and enhance the C₂₊ species formation during CO₂ hydrogenation.⁴²

So far, the effect of reducible supports on bimetallic Ni,Fe catalysts for CO₂ hydrogenation has not been studied to the same extent as the support effect on monometallic catalysts or the Fe promotion effect on unreducible supports.⁴³ In this

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work, we first confirmed the intrinsic optimal Ni : Fe ratio for methanation on non-reducible, but basic MgO. Then, monometallic Ni and bimetallic Ni,Fe catalysts at this optimal ratio supported on MgO, ZrO₂ and CeO₂ have been compared to study if the promotional effect of Fe depends on the varying reducibility of the support by kinetic measurements, H₂-TPR, CO₂-TPD, and *in situ* DRIFTS measurements.

2. Experimental section

2.1. Catalyst preparation

The catalysts were prepared by the incipient wetness impregnation method. For the support of MgO, 1.25 mL of an aqueous solution containing Fe(NO₃)₃·9H₂O (99.5%, Grüssing GmbH) and Ni(NO₃)₂·6H₂O (99.9%, abcr GmbH) was added dropwise to 1 g of MgO (>99%, Alfa Aesar). The metal salt concentration was adjusted to reach a metal total (Ni and/or Fe) loading of 5 wt% on MgO. The wet powder was further stirred and then treated by ultrasonication for 15 min. After drying at 80 °C overnight, the collected pre-catalysts were directly reduced with 10% H₂/N₂ at 450 °C for 2 h prior to the catalytic tests. By changing the ratio between Fe(NO₃)₃·9H₂O and Ni(NO₃)₂·6H₂O in the aqueous solution, the molar ratio of Ni : Fe has been adjusted to 1 : 0, 5 : 1, 2 : 1, 1 : 1, 1 : 2, and 0 : 1. The reduced samples are labelled as Ni/MgO, Ni,Fe(83 : 17)/MgO, Ni,Fe(67 : 33)/MgO, Ni,Fe(50 : 50)/MgO, Ni,Fe(33 : 67)/MgO, and Fe/MgO respectively. The same procedure has been applied to the supports of ZrO₂ (99%, Alfa Aesar) and CeO₂ (99.5%, Alfa Aesar), but only monometallic Ni and a Ni : Fe ratio of 5 : 1 catalysts were selected, resulting in samples labelled as Ni/ZrO₂, Ni,Fe(83 : 17)/ZrO₂, Ni/CeO₂ and Ni,Fe(83 : 17)/CeO₂.

2.2. Catalysts characterization

Inductively coupled plasma optical emission spectroscopy (ICP-OES) was applied to analyze the elemental compositions of those catalysts supported on MgO and CeO₂. The pre-catalysts were dissolved in a mixture of nitric acid and hydrogen peroxide and measured on an Avio 200 ICP OES equipped with a S23 Autosampler (PerkinElmer). Due to the low solubility of ZrO₂, the metal compositions of these catalysts were determined in a scanning electron microscope by energy dispersive X-ray spectroscopy (SEM-EDX) using a Gemini Ultra55Plus (Zeiss).

The surface area was measured by N₂ physisorption using the Brunauer–Emmett–Teller (BET) method on a BELSORP MAX (Microtrac Retsch GmbH). The samples were pre-treated at 100 °C for 2 h in vacuum to remove adsorbed moisture and air. The isotherm was measured at –197 °C in liquid nitrogen.

Powder X-ray diffraction (PXRD) was used to study the crystal structure of the catalysts using Mo K α radiation on a STADI P diffractometer (STOE). To avoid the re-oxidation of the reduced catalysts, the reduced samples were prepared in an Ar glovebox as sealed capillaries with 0.7 mm diameter and later measured in the Debye–Scherrer mode.

Hydrogen temperature programmed reduction (H₂-TPR) was performed on a BELCAT II catalyst analyzer (Microtrac Retsch GmbH) to study the reducibility of the catalysts. The measurements were conducted with 50 mg of the samples filled in a quartz glass fixed-bed micro-reactor. After pretreatment at 120 °C under on Ar flow for 1 h, the reactor was cooled down to 40 °C and then heated to 1000 °C at a rate of 6 °C min^{–1} in a 50 sccm of 10% H₂/Ar flow. The reactor was later kept at 1000 °C for 15 min before cooling down and the H₂ consumption was measured with a thermal conductivity detector (TCD). The degree of the reduction (%) of the supported metal was estimated by the H₂ consumption assuming that full reduction corresponds to a state with all Ni and Fe being metallic, CuO was used to calibrate the H₂ consumption.

CO₂ temperature programmed desorption (CO₂-TPD) profiles were recorded with the home-built reaction setup equipped with TCDs on the inlet and outlet of the quartz tube reactor. Prior the measurement, all samples were heated to 450 °C with a rate of 5 °C min^{–1} in a 10 mL min^{–1} of 20% H₂/Ar flow. After an isothermal reduction at 450 °C for 2 h, samples were cooled down to 50 °C in a flow of He to remove the residual gases from the reduction. CO₂ adsorption was carried out at 50 °C for 2 h with a flow rate of 10 mL min^{–1}. Then a 10 mL min^{–1} of He was introduced to remove physically adsorbed CO₂ before the desorption. CO₂-TPD curve was drawn at the temperature range from 50 to 600 °C with the heating rate of 10 °C min^{–1}. The number of surface basic sites was calculated using the integration of the calibrated CO₂ peak below the curve. The raw data were smoothed first and then fitted (Fig. S3, ESI†) into three Gaussian peaks at different temperature ranges representing the weak, medium and strong basic sites.

Transmission Electron Microscopy (TEM) was employed to evaluate the nanoparticle size of the freshly reduced samples. TEM images were taken in the bright field mode on a Tecnai F30 G2 S-TWIN (Thermo Fisher Scientific) equipped with a field emission gun (FEG) using an acceleration voltage of 300 kV and a spherical aberration coefficient of 1.2 mm. Due to the low metal loading and small nanoparticle size, the brightness and contrast of TEM images were adjusted to clearly identify metal particles.

Diffuse Reflectance Infrared Fourier Transform Spectroscopy (DRIFTS) measurements were performed on the reduced catalyst following the adsorption and desorption of CO used as a probe molecule. The samples were heated up to 450 °C with 5 °C min^{–1} in an 80 mL min^{–1} flow of 10% H₂/He and then an isothermal reduction was performed for 2 h. Afterwards, the reduced catalysts were purged in pure He for 30 min and then cooled down to 10 °C. The analysis was performed under 80 mL min^{–1} of 0.5% CO/He mixture flow for 40 minutes during adsorption and pure He flow for 40 minutes during desorption. The spectra were collected by a iS50 FTIR spectrometer (Thermo Fisher Scientific) equipped with a high-temperature reaction chamber (Praying Mantis, Harrick) coupled with a dome mounting ZnSe windows. All



measurements were collected by the OMNIC software *via* an automatic series collection using a Gram-Schmidt method. The temperature was controlled by Harrick ATC and a Huber Ministat 125 chiller. No compensation of H₂O and CO₂ signal was used.

2.3. Kinetic measurements

Catalytic tests were performed in a U-shaped micro-reactor where the catalyst bed was fixed by quartz wool plugs. A thermocouple was inserted directly into the catalyst bed to measure the temperature. After isothermal reduction at 450 °C for 2 h, the reactor was cooled down to 350 °C and 50 sccm of a reaction gas mixture was injected into the reactor to start the CO₂ hydrogenation test at atmospheric pressure, consisting of 9.6% CO₂/N₂, 38.4% H₂ and balanced N₂. The inlet gas flow was controlled by mass flow controllers (Bronkhorst) and the gas composition at the exhaust was evaluated by a micro-GC (INFICON). In a typical catalytic test, 30 mg of MgO-supported catalysts were used while 10 and 3 mg of the ZrO₂ and CeO₂-supported catalysts were sufficient to reach a similar CO₂ conversion of less than 25% to ensure the measured selectivity for methanation can be compared at similar conversion levels. The catalyst sieve fraction of 200–300 μm was diluted with 350–460 μm SiC (Thermal Scientific) to give a total weight of 200 mg to avoid heat and mass transfer issues.

The activation energies (E_a) of selected catalysts were measured at temperatures ranging from 315 to 355 °C and reaction orders for H₂ and CO₂ were tested in the same reactor at 350 °C by changing the respective partial pressures. All the kinetic measurements were conducted under low CO₂ conversion of less than 25%.

As shown in eqn (1), the reaction rates of CO₂ hydrogenation were defined as the CH₄ production rate (r_{CH_4}) for methanation and the CO production rate (r_{CO}) for the RWGS normalized to the weight of active metals. The conversion of CO₂ and the selectivity for CH₄/CO were calculated as eqn (2) and (3).

$$r_{\text{CH}_4/\text{CO}} = \frac{\text{Flow rate} \times \text{conc.}_{\text{CH}_4/\text{CO}}}{g_{\text{cat.}} \times \text{wt}\%} \quad (1)$$

$$\text{Conv.}_{\text{CH}_4/\text{CO}} = 1 - \frac{\text{conc.}_{\text{CO}_2,\text{out}}}{\text{conc.}_{\text{CO}_2,\text{in}}} \quad (2)$$

$$\text{Selec.}_{\text{CH}_4/\text{CO}} = \frac{\text{conc.}_{\text{CH}_4/\text{CO}}}{\text{conc.}_{\text{CH}_4/\text{CO}} + \text{conc.}_{\text{CO}/\text{CH}_4}} \quad (3)$$

3. Results

The metal loading of the as-synthesized samples was checked by ICP-OES and SEM-EDX, which confirmed the total metal loading are approximately 5 wt% for all the samples (Tables S1 and 2, ESI[†]). For the bimetallic catalysts, also the measured Ni:Fe ratios are in reasonable agreements with the nominal values (Tables S1 and 2, ESI[†]). The BET surface areas of the commercial supports are approximately 15 (MgO), 26 (ZrO₂), and 64 (CeO₂) m² g⁻¹ (Table S3, ESI[†]). The reducibility of the

Ni,(Fe)/MgO catalysts was studied by H₂-TPR (Fig. 1a). With increasing Fe content, the reduction peaks gradually shift to higher temperatures. Such trend is consistent with previous reports and can be explained by the less noble nature of iron compared to nickel.^{44,45} A single TPR peak was observed for the reduction into the metallic state indicating presence of uniform metal species.^{46,47} Only a weak shoulder is observed at the lower temperature side of the TPR profile (around 400 °C) with increasing Fe content, which can be assigned to the intermediate reduction of Fe³⁺ to Fe²⁺.^{47–49} There is no significant reduction at temperatures higher than 500 °C for all the MgO-supported catalysts confirming the unreducible nature of the support at these conditions. Thus, an isothermal reduction at 450 °C for 2 h was chosen to reach a full reduction of both Ni and Fe. The XRD patterns of the freshly reduced samples are shown in Fig. 1b. The major crystalline phase in all catalysts is MgO while there are no obvious reflections of Ni or Fe metal detected, which can be explained with the low loading of active metal and the small crystallite size.

The catalytic performance of the reduced Ni,Fe/MgO catalysts for CO₂ hydrogenation was evaluated at 350 °C for 16 h after the above-described reduction step. Both of the monometallic Ni and Fe catalysts showed stable activity in methanation and RWGS (Fig. 2) and, as expected, Ni had a higher activity and selectivity to methane. Note that it was reported that the addition of low amounts of Fe to Ni catalysts can prohibit the deactivation of monometallic Ni catalyst,^{22,50} but no such deactivation was present for the MgO-supported monometallic Ni in our experiment. However, the bimetallic Ni,Fe catalysts showed high CO₂ conversions at the beginning, but suffer from gradual deactivation over 16 h time on stream (TOS). In this process, methanation deactivated stronger than RWGS leading to a decrease of CH₄ selectivity (Fig. 2b–d). Sintering and carbon deposition are the most probable reasons for such initial deactivation.^{51–53} Under methanation reaction conditions, it has been reported that small Ni nanoparticles may coalesce to form bigger particles, followed by the Ostwald ripening.^{54,55} The particle size growth can also be associated to the formation of metal carbonyls due to the enhanced mobility of Ni(CO)_x species.^{55–57} The particle size increase will result in the loss of surface area and active sites, eventually resulting in the deactivation of catalysts. After 16 h TOS, both the methanation and RWGS rates have approached to a steady-state for each catalyst, revealing a volcano-shaped trend with the Fe loading (Fig. 3). The optimal Ni:Fe ratio for the CO₂ hydrogenation was identified as 5:1, *i.e.* Ni,Fe(83:17)/MgO, with the highest activity in both methanation (0.035 mmol_{CH₄} g_{metal}⁻¹ s⁻¹) and RWGS (0.302 mmol_{CO} g_{metal}⁻¹ s⁻¹). It was proposed that iron oxide species can be formed from the oxidation of iron by CO₂ and H₂O in the gas stream. These can be dispersed on the Ni particles, and the redox cycle of oxidized Fe species can promote the CO₂ hydrogenation.^{20–23} A similar dynamic promotion effect in Co,Fe bimetallic catalysts in CO₂ hydrogenation was interpreted as a result of alloy formation after reduction of bimetallic catalysts.⁵⁸ In both scenarios, a too strong increase in Fe loading will cause either blocking or



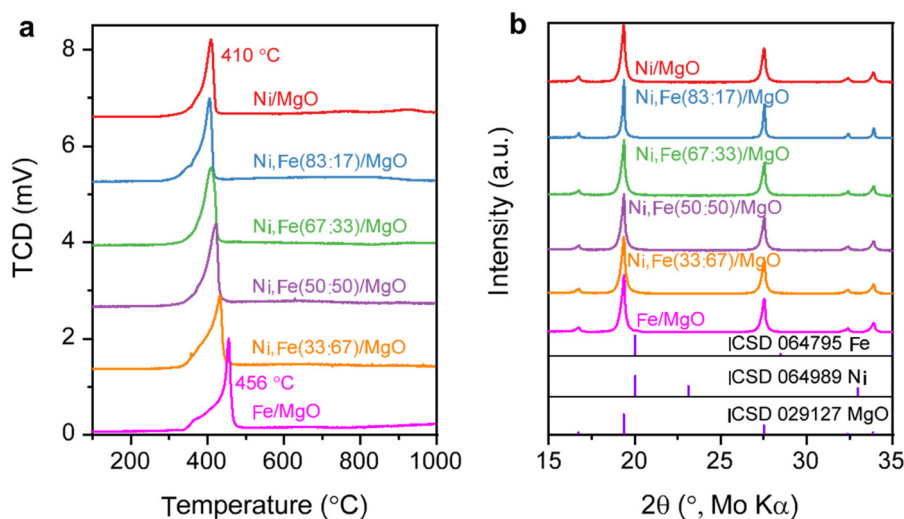


Fig. 1 (a) H₂-TPR profile for Ni,Fe monometallic and bimetallic catalysts supported on MgO and (b) their PXRD patterns after isothermal reduction of 450 °C.

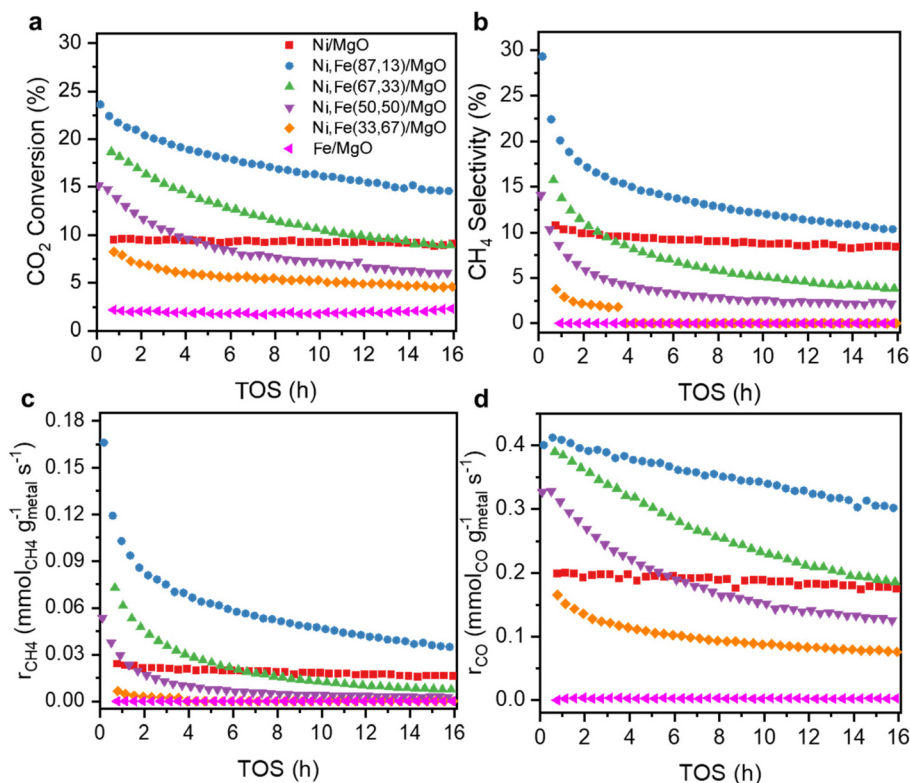


Fig. 2 (a) CO₂ conversion, (b) selectivity towards CH₄, (c) methanation rate and (d) RWGS rate as a function of TOS for MgO-supported catalysts with various Ni : Fe ratios at 350 °C.

substitution of active Ni sites at the surface, which explains the decreased activity with higher Fe loading, which is not uncommon for promoter species. Besides, our Ni,Fe bimetallic catalysts also suffer from deactivation, which might be associated with the gradual and dynamic formation of such oxidized Fe “islands”.

The Ni,Fe(83:17)/MgO catalyst reached the highest reaction rate of both methanation and RWGS reactions with the highest CH₄ selectivity in this series of catalysts (~10.2%) (Fig. 2b–d). This ratio of 5 : 1 for Ni : Fe has been employed in the study of the other two supports, *i.e.* ZrO₂ and CeO₂, to investigate the support effect on the Ni,Fe bimetallic catalysts for CO₂ hydrogenation.



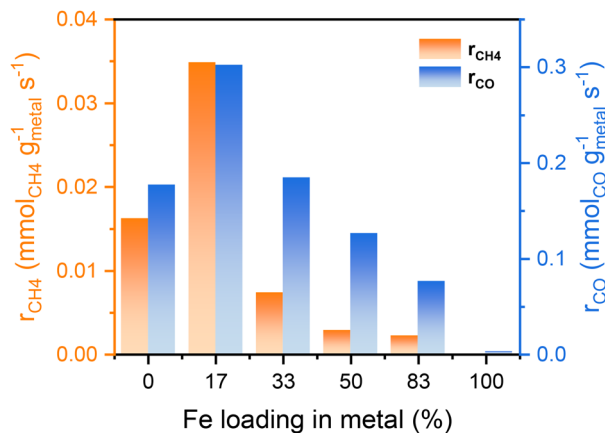


Fig. 3 Reaction rates of both methanation and RWGS after 16 h test for different Fe loadings in Ni,(Fe)/MgO catalysts.

The reducibility of these two series catalysts was studied first. The H₂-TPR profiles of ZrO₂ and CeO₂ supported catalysts exhibit two reduction peaks below 400 °C, which could result from weakly and strongly bound Ni²⁺ and Fe³⁺ species^{59,60} (Fig. 4a), indicating the reduction of the metal completes before 450 °C. Besides, the reduction degree (%) of the supported metal estimated by the H₂ consumption assuming that all Ni and Fe can be fully reduced were compared among these mono- and bimetallic supported catalysts, following the order CeO₂ → ZrO₂ → MgO-supported catalysts (Fig. 4c). The higher reduction degree of CeO₂- and ZrO₂-supported catalysts agree with the reducibility of these oxides. For the CeO₂-supported catalysts, reduction degree even higher than 100% clearly indicate that support species have been reduced together with the metal oxide species at low temperature (<300 °C), agreeing well with the previous report of co-reduction of highly dispersed metal and support-CeO₂,^{61–63} which was explained by metal-

support interaction (MSI) and/or hydrogen spillover effects.^{64–68}

The XRD patterns of these reduced samples show broad reflections for Ni and Fe at around 20° 2θ, which might be due to the small metal particle size or low metal loading (Fig. 4b). The broad XRD peak profiles agree well with the TEM images showing that most metal nanoparticles have a size around 5–10 nm (Fig. 5). Similar nanoparticle sizes were found for all six catalysts, which indicates that the comparison here is not strongly affected by possible size effects on CO₂ hydrogenation.⁶⁹

All supported Ni and Ni,Fe(83 : 17) catalysts were studied in the CO₂ hydrogenation reaction at 350 °C for 16 h. The CO₂ conversion is similar for monometallic Ni and bimetallic Ni,Fe (83 : 17) catalysts on both supports ZrO₂ and CeO₂, suggesting in contrast to MgO-supported catalysts, no clear promotion effect of Fe on the CO₂ conversion for these reducible supports (Fig. S2, ESI[†]). Monometallic Ni even showed a higher CO₂ methanation rate than bimetallic Ni,Fe(83 : 17) if supported on ZrO₂ or CeO₂ (Fig. 6) showing a negative effect of Fe addition on the methane selectivity. Actually, Winter *et al.* already found that the introduction of Fe to Ni catalysts supported on CeO₂ will increase the selectivity towards RWGS.⁷⁰ They proposed that the formation of oxidized Fe will weaken the CO binding strength, resulting in desorption. Since the adsorbed CO is one of the potential intermediates for methanation, this can explain the lower selectivity to CH₄ in favor of RWGS.⁷¹ However, this is very different to MgO-supported catalysts where the addition of Fe enhances both CO₂ conversion and methane selectivity, though on a lower level of around 10%. The absolute selectivity to methane for Ni/CeO₂ and Ni/ZrO₂ were much higher than on MgO and reached 82.3% and even 94.4%, indicating that the methanation reaction is greatly promoted by the reducible nature of the oxide support with a much stronger effect than the presence or absence of Fe.

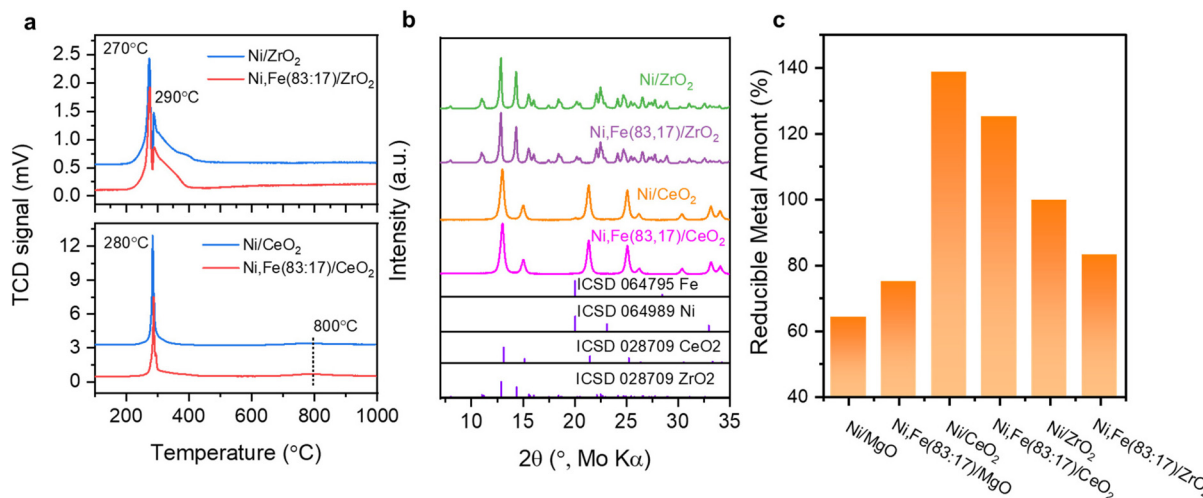


Fig. 4 (a) H₂-TPR profile for Ni monometallic and Ni,Fe bimetallic catalysts supported on ZrO₂, CeO₂; (b) the PXRD patterns of the samples reduced isothermally at 450 °C and (c) the estimated degree of reduction, where valued beyond 100% refer to co-reduction of support species.



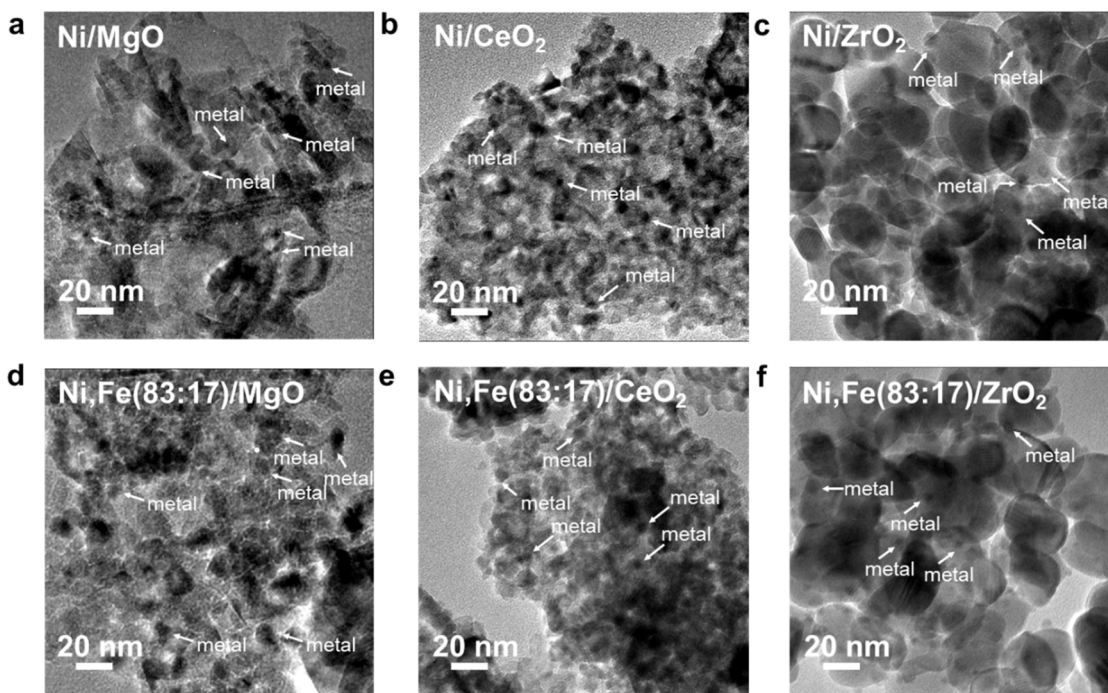


Fig. 5 TEM images for Ni–Fe monometallic and bimetallic catalysts supported on (a) & (d) MgO, (b) & (e) ZrO₂ and (c) & (f) CeO₂.

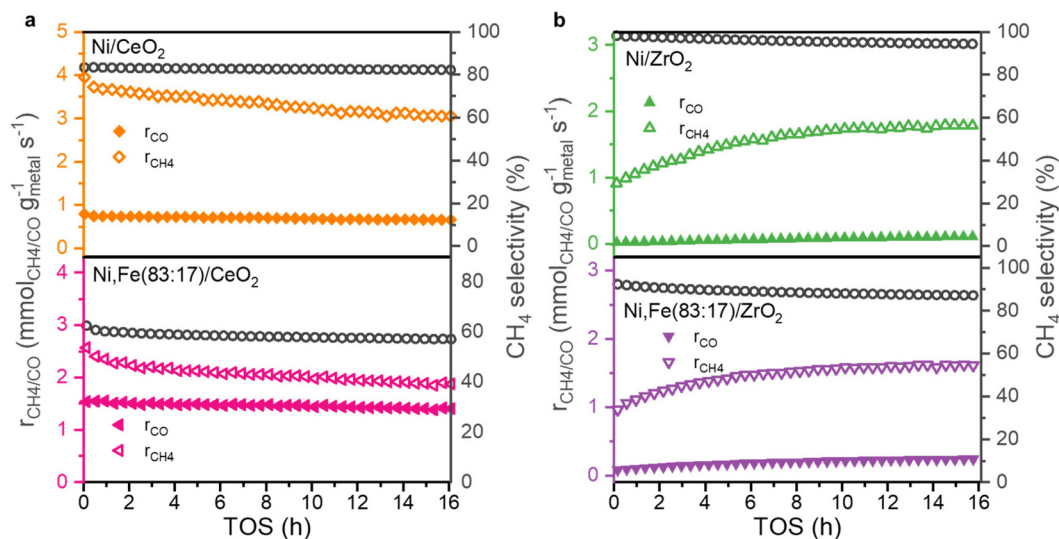


Fig. 6 Methanation and RWGS rates and selectivity towards CH₄ as a function of TOS for (a) CeO₂ and (b) ZrO₂-supported monometallic and bi-metallic catalysts for 16 h TOS at 350 °C.

While the selectivity remains quite stable during the 16 hours TOS (Fig. 6), the conversions show an interesting trend (Fig. S2, ESI†). The CeO₂-supported catalysts show a slight deactivation, which is much less pronounced compared to Ni,Fe(83:17)/MgO, but for the ZrO₂ supports, a clear activation phase in both of methanation and RWGS reactions was observed. A similar behavior was found before in Ru/ZrO₂ in CO_x methanation,^{39,72} which was explained by the slow

reduction of ZrO₂ leading to increasing metal-support interaction during the hydrogenation reaction.

The reaction pathways on Ni/ZrO₂ and Ni/CeO₂ catalysts discussed in the literature indeed assume a strong involvement of the support. CO₂ is first adsorbed on a hydroxyl group of the ZrO₂/CeO₂ support. There, it can be hydrogenated to formate by dissociated hydrogen from metallic Ni under desorption of H₂O.^{73–77} The formate intermediate can either decompose and



then transfer onto the surface of the metallic Ni nanoparticle or directly be converted on the $\text{ZrO}_2/\text{CeO}_2$ surface to CH_4 by combing with more dissociated hydrogen spilled over from Ni.^{73,75–77} Both reaction pathways require a strong participation of $\text{ZrO}_2/\text{CeO}_2$ in the CO_2 activation steps. The latter pathway even avoids adsorption of any carbon- or oxygen-bound intermediates on the surface of Ni, which could help to keep the metallic Ni particles free from deactivating species. Contrarily, for the catalysts supported on MgO, CO_2 and its CO^* or HCOO^* intermediates may absorb mainly on the Ni surface, which increases the probability of strong deactivation by oxidation or coking.⁷

Comparing the steady-state activity (after 16 h TOS) among these Ni and Ni,Fe catalysts supported on three different oxides, the CeO_2 supported catalysts show the highest activity. Ni/ CeO_2 has the highest methanation rate of $3.0 \text{ mmol}_{\text{CH}_4}$,

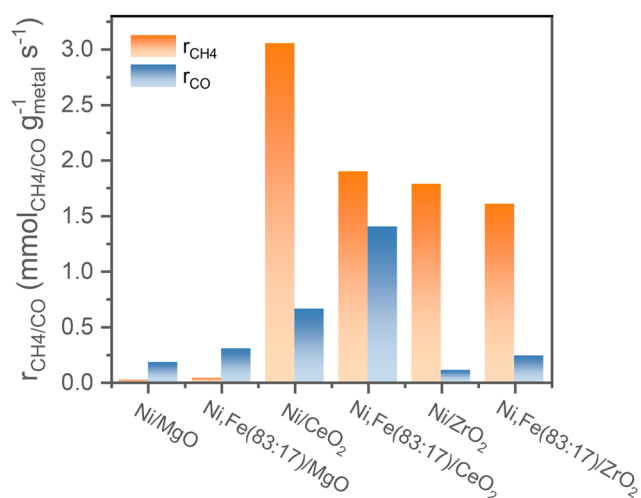


Fig. 7 Comparison of the reaction rates of the different supported mono- or bimetallic catalysts for both methanation and RWGS after 16 h TOS.

$\text{g}_{\text{metal}}^{-1} \text{ s}^{-1}$ and Ni,Fe(83 : 17)/ CeO_2 has the highest RWGS rate of $1.3 \text{ mmol}_{\text{CO}} \text{ g}_{\text{metal}}^{-1} \text{ s}^{-1}$ (Fig. 7). The steady-state activity of the methanation reaction followed the order of MgO-supported < ZrO_2 -supported < CeO_2 -supported, which agrees well with the reducibility order: $\text{MgO} < \text{ZrO}_2 < \text{CeO}_2$ and underlines to role of a SMSI or electronic metal-support interaction (EMSI) effect.^{18,33,39,78} The basicity of the Ni-based monometallic catalysts and their bare support were investigated through CO_2 -TPD experiments (Fig. 8 and Table 1). CO_2 -TPD profiles revealed three Gaussian peaks representing the weak (100–170 °C), medium (222–302 °C), and strong (396–501 °C) basic sites. The low-temperature peak corresponds to the CO_2 desorption from weak Brønsted sites.^{79–82} The medium-temperature peak originates from bidentate carbonates on metal-oxygen pairs such as Mg–O.^{79–82} The high-temperature peak is correlated to the desorption of unidentate carbonates on low-coordination O^{2-} anions.^{79–83} Based on the quantitative analysis of CO_2 -TPD profiles, Ni/ CeO_2 catalysts have the highest content of weak basic sites ($29 \mu\text{mol}_{\text{CO}_2} \text{ g}_{\text{cat}}^{-1}$) among these three supported Ni catalysts, which indicates that the CeO_2 support can facilitate the CO_2 adsorption, contributing to its high activity.^{84–86} Furthermore, compared to the bare support, our results suggest that CeO_2 -supported catalysts have a higher total density of basic sites ($85 \mu\text{mol}_{\text{CO}_2} \text{ g}_{\text{cat}}^{-1}$ for Ni/ CeO_2 vs. $54 \mu\text{mol}_{\text{CO}_2} \text{ g}_{\text{cat}}^{-1}$ for CeO_2), whereas a lower total basic sites

Table 1 Calculated basic site density based on the results from CO_2 -TPD of the bare supports and the reduced catalysts

	Weak [$\mu\text{mol g}^{-1}$]	Medium [$\mu\text{mol g}^{-1}$]	Strong [$\mu\text{mol g}^{-1}$]	Total [$\mu\text{mol g}^{-1}$]
MgO	15	87	65	167
Ni/MgO	17	72	21	111
ZrO ₂	11	52	22	85
Ni/ZrO ₂	12	7	7	26
CeO ₂	27	8	19	54
Ni/CeO ₂	29	19	37	85

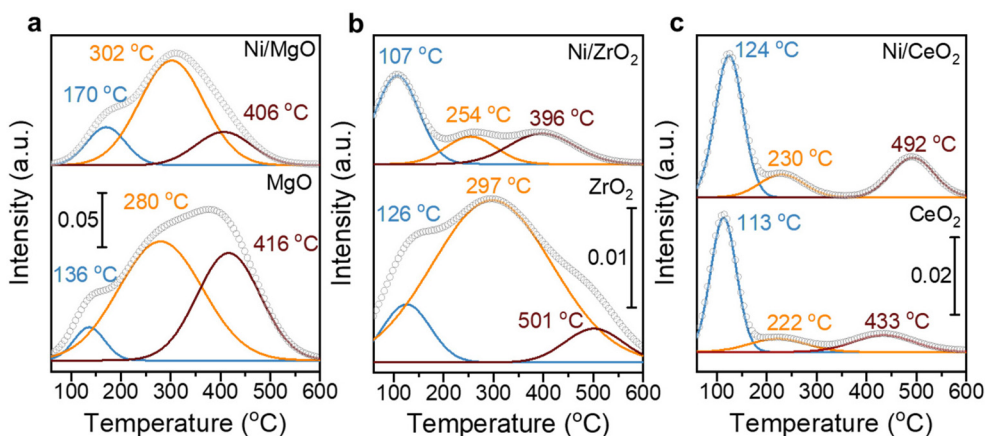


Fig. 8 CO_2 -TPD profiles at 10 °C min^{-1} for (a) MgO, (b) ZrO_2 , (c) CeO_2 supports, and their corresponding Ni loaded catalysts where grey dots represent the smoothed raw profiles and the solid color lines indicate the fitting results for weak, medium and strong basic sites at different temperatures.



density was observed for the MgO-supported ($111 \mu\text{mol}_{\text{CO}_2} \text{g}_{\text{cat}}^{-1}$ for Ni/MgO vs. $167 \mu\text{mol}_{\text{CO}_2} \text{g}_{\text{cat}}^{-1}$ for MgO) and ZrO₂-supported ($26 \mu\text{mol}_{\text{CO}_2} \text{g}_{\text{cat}}^{-1}$ for Ni/ZrO₂ vs. $85 \mu\text{mol}_{\text{CO}_2} \text{g}_{\text{cat}}^{-1}$ for ZrO₂) samples. Notably, experimental results have revealed that MgO and its catalysts release significantly more CO₂ during the TPD experiments. This increased CO₂ release is likely related to the presence of bulk carbonates formed during CO₂ adsorption^{87,88} and is not necessarily indicative of surface basicity. Nevertheless, these results clearly demonstrate that the SMSI in Ni/CeO₂ catalyst results in interfacial sites more favorable for CO₂ activation/conversion. As Ni/CeO₂ contains a higher density of basic sites than the bare support CeO₂, and given that the density of basic sites can be correlated to the oxygen vacancy density, this is consistent with the results of the H₂-TPR analysis.^{89,90}

Further kinetic experiments were conducted for Ni and Ni, Fe catalysts to determine the apparent activation energy (E_a) and reaction order of CO₂ and H₂ with respect to methanation and RWGS for a deeper investigation of the support effect. From the Arrhenius plot shown (Fig. 9a–c), the apparent activation energy E_a for MgO and CeO₂ supported catalysts are very similar between 71 and 79 kJ mol⁻¹ in both methanation and RWGS reactions. However, the intercepts in the Arrhenius plot for the MgO-supported catalysts are much lower than for the CeO₂-supported catalysts, suggesting more active sites from the latter catalysts. This interpretation agrees with the previous hypothesis that the metal-support interaction provides additional abundant active sites at the ceria-metal interfaces. Besides, the high reducibility of CeO₂ results into more oxygen vacancies introducing the so-called electronic metal-support-interaction (EMSI) effect, which can activate CO₂ at lower temperatures.¹⁸ For ZrO₂-supported catalysts, the Arrhenius plots suggest a different activation energy compared to MgO- and CeO₂-supported catalysts. Here, the E_a of methanation is 94.4 kJ mol⁻¹, which is almost half of that of RWGS (196.6 kJ mol⁻¹), implying that the relative barrier energy for RWGS is very high

on this type of catalyst. The big difference in activation energy for these two reactions on Ni/ZrO₂ can account for the highest methane selectivity for ZrO₂-supported catalysts.

The reaction orders of CO₂ and H₂ for monometallic Ni and bimetallic Ni,Fe catalysts were studied for understanding of the limiting factors of reactants activation (Fig. 10 and 11). For the three monometallic Ni catalysts, an increasing in CO₂ partial pressure affected the RWGS more than methanation, suggesting that the RWGS is limited by the activation of CO₂, especially for Ni/CeO₂ with the highest CO₂ reaction order of 1.08 in RWGS (Fig. 10a–c). At the same time, the H₂ partial pressure variation affected the methanation rate stronger and positively showing that the methanation is rather limited by the activation of H₂. The RWGS was even affected negatively by an increase in the H₂ partial pressure for the two catalysts on the reducible supports, CeO₂ and ZrO₂, especially for ZrO₂-supported catalysts with a value of -0.86 (Fig. 10d–f). The high selectivity towards CH₄ for the ZrO₂-supported catalysts can therefore also be explained by the hydrogen-rich condition (H₂:CO₂ = 4:1), which also agrees well with the previous results of activation energy. In addition, as the calculated reduction degree of CeO₂ supported catalysts exceeds 100%, indicating that not only Ni and Fe were activated to the metallic state, but CeO₂ was also partially reduced to CeO_{2-x}, forming surface oxygen vacancies.^{91,92} The oxygen vacancies donate electrons which could participate in the adsorption and activation of CO₂, explaining the reason that the CeO₂ supported catalysts with relatively rich oxygen vacancies have a higher CO₂ reaction order in both methanation and RWGS (Fig. 10). Furthermore, the H₂-spillover effect is supported by the H₂-TPR result where the reduction of surface CeO₂ occurs at around 280 °C, significantly lower than the surface reduction temperature of pure CeO₂, which is normally around 450 °C.^{62,93,94} This indicates that the dissociated H migrate from the metal to the metal-support interface and can hydrogenate the activated CO₂.

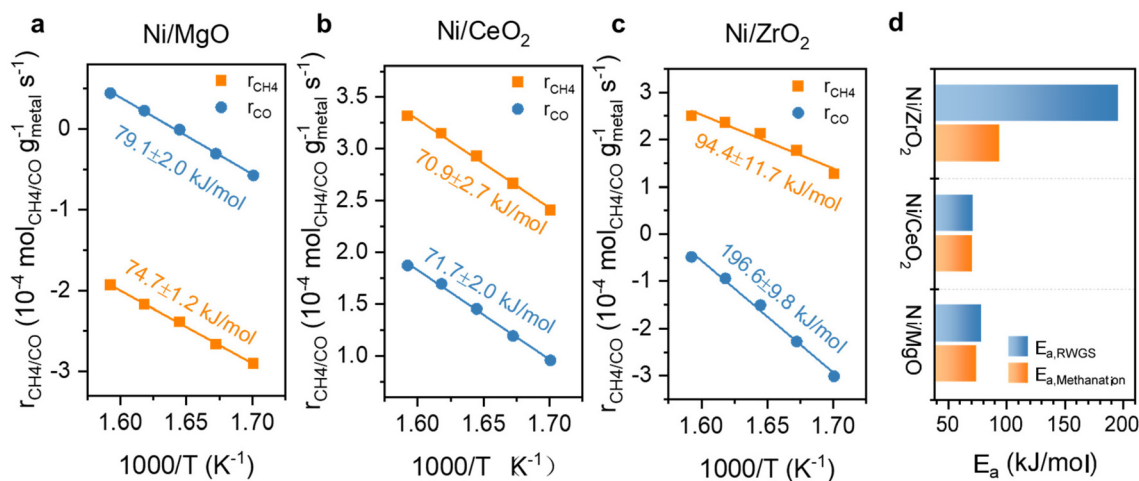


Fig. 9 Arrhenius plots with calculated E_a for Ni catalysts supported on (a) MgO, (b) CeO₂, (c) ZrO₂ and (d) comparison of their E_a s for RWGS and methanation.



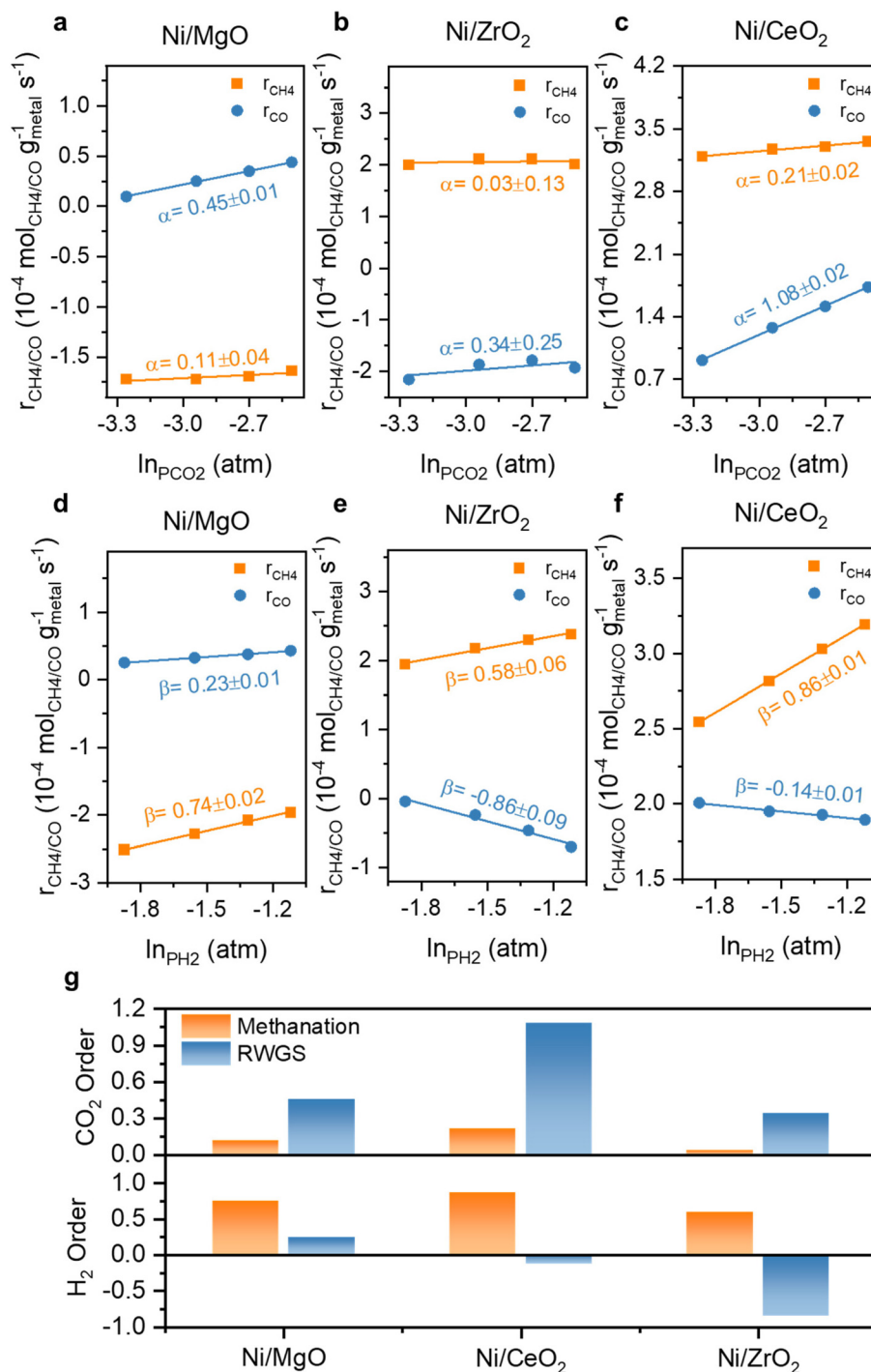


Fig. 10 The reaction order of CO₂ (α) and H₂ (β) at 350 °C for Ni monometallic catalysts supported on (a) & (d) MgO, (b) & (e) ZrO₂ and (c) & (f) CeO₂ and (g) their comparison.

The reaction orders of CO₂ and H₂ for the bimetallic Ni,Fe catalysts are shown in Fig. 11. Compared to the monometallic catalysts, there was a more significant increase in both reaction orders of CO₂ and H₂ for methanation compared to RWGS reaction on the MgO-supported catalyst, corresponding to the higher methanation selectivity on the bimetallic catalysts (Fig. 2). For the ZrO₂-supported catalyst Ni,Fe(83 : 17)/ZrO₂, the H₂ reaction order

of -0.21 in RWGS was still negative but has increased strongly from -0.86 observed in Ni/ZrO₂. This less negative H₂ reaction order showed that the H₂-inhibition of RWGS was mitigated compared to the monometallic Ni catalyst, consistent with a lower RWGS selectivity on this bimetallic catalyst (Fig. 6b). Notably, introducing Fe to Ni/CeO₂ minorly modified the reaction order of CO₂ or H₂ in both methanation and RWGS.



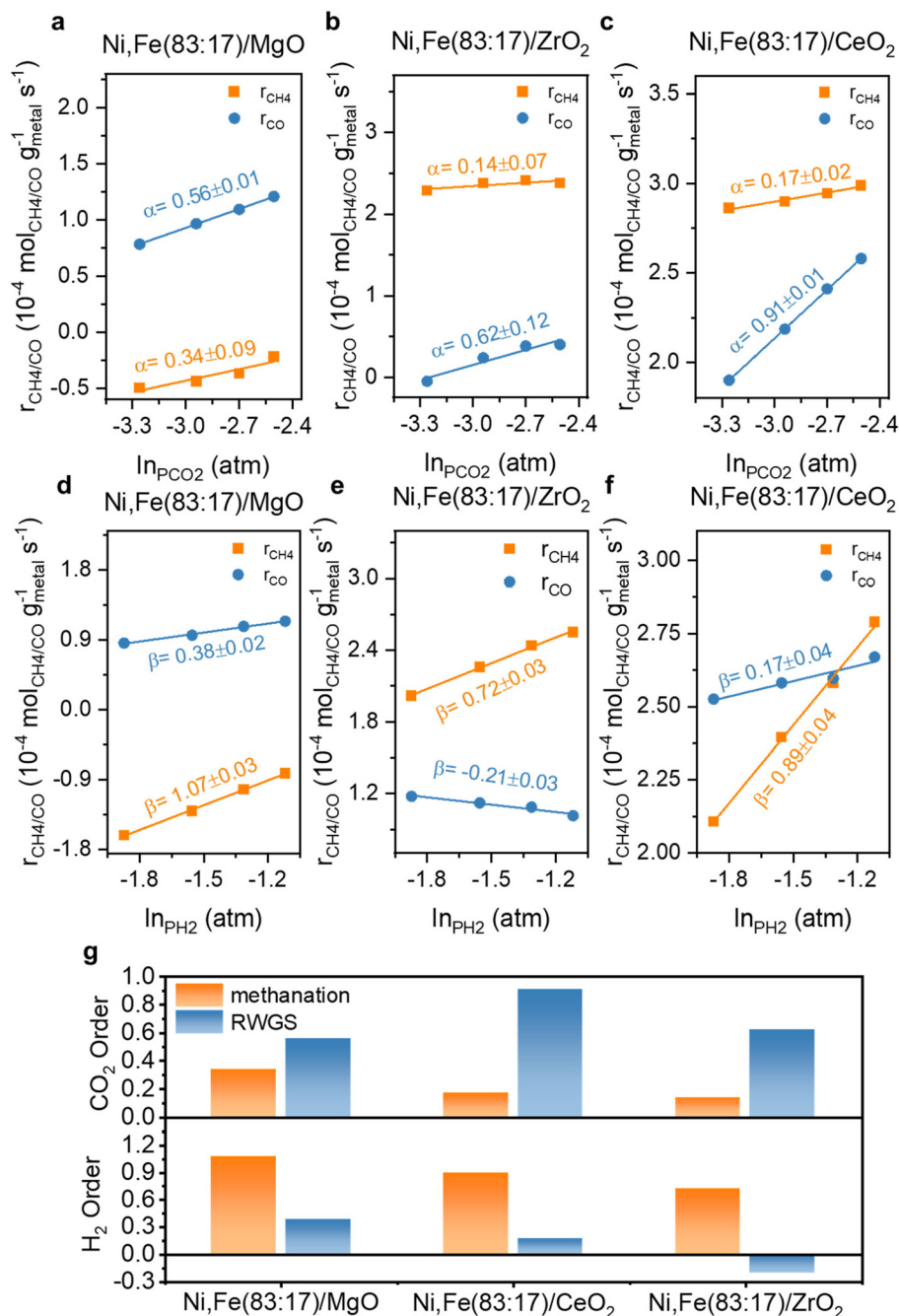


Fig. 11 The reaction order of CO₂ (α) and H₂ (β) at 350 °C for Ni,Fe bimetallic catalysts supported on (a) & (d) MgO, (b) & (e) ZrO₂ and (c) & (f) CeO₂ and (g) their comparison.

Additionally, *in situ* DRIFTS combined with CO as a probe were performed on the Ni monometallic catalysts to study the impact of the support on the surface properties of Ni. The formed surface species during CO adsorption can be divided in two main regions: one with spectral features above 1800 cm⁻¹ and the other with features below 1800 cm⁻¹. For Ni/MgO, the spectra region above 1800 cm⁻¹ is characterized by adsorbed CO (CO_{ad}) species *i.e.* the metal carbonyl, in particular the linear Ni⁰-CO species (2056 and 2029 cm⁻¹) and the

bridged (Ni⁰)_x-CO species (1945 and 1919 cm⁻¹).^{95,96} At features below 1800 cm⁻¹, several bands can be identified as carbonate, bicarbonate, and formate species.^{95,97-99} The intensity of carbonaceous/carbonate-like species and CO_{ad} species increased as a function of the CO adsorption time (Fig. 12a). In the desorption period (Fig. S4, ESI[†]), CO_{ad} species gradually diminished and left a bit amount of CO_{ad} adsorbed on the Ni surface while carbonate, bicarbonate, formate species region remained after 40 min of desorption.



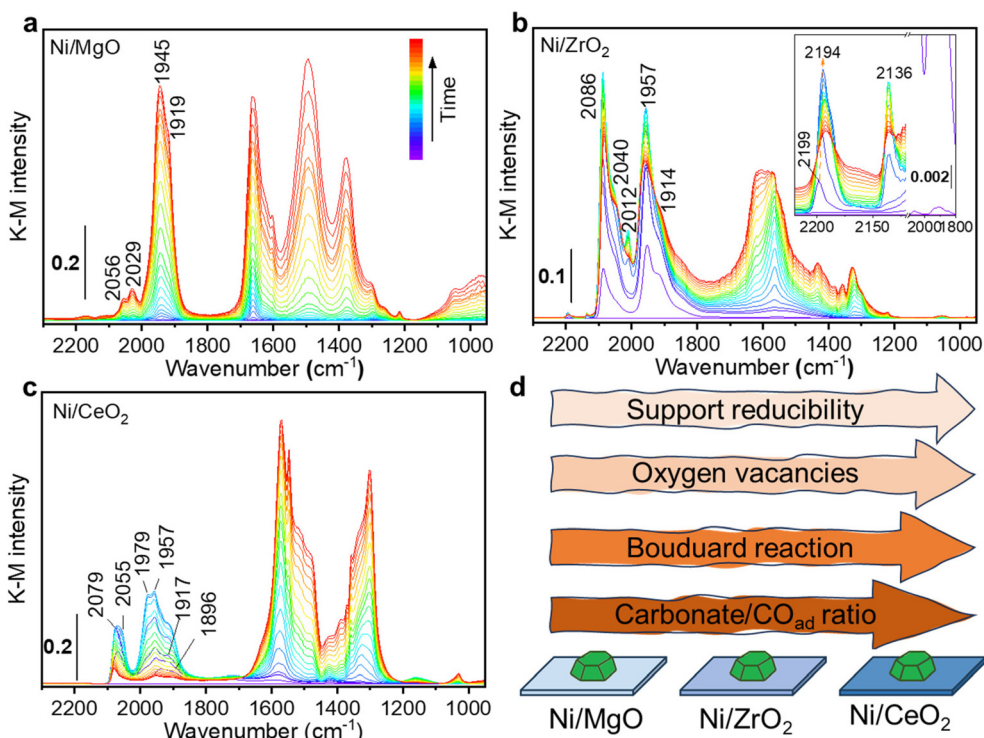


Fig. 12 *In situ* DRIFTS spectra of (a) Ni/MgO, (b) Ni/ZrO₂ and (c) Ni/CeO₂ catalysts collected during CO adsorption, inset images in (b) representing the zoom-in spectra at features above 2100 cm⁻¹. Illustration (d) of the structural properties changes in three different supported Ni Catalysts.

For Ni/ZrO₂, CO adsorption resulted in the formation of a very intense peak at 2086 cm⁻¹, with a shoulder at 2040 cm⁻¹, which was ascribed to linear Ni⁰-CO (Fig. 12b). At higher frequency, two tiny peaks were identified and ascribed to linear Ni²⁺-CO (2199–2194 cm⁻¹) and Ni⁺-CO (2136 cm⁻¹).⁹⁶ Besides, the two bands at 1957 and 1914 cm⁻¹ were ascribed to bridged (Ni⁰)_x-CO species. In the lower frequency region below 1800 cm⁻¹, several bands were observed and identified as carbonate and bicarbonate species, similar to Ni/MgO. During the CO adsorption, the intensity of all CO_{ad} bands increased to a maximum after about 6 minutes and then started to decrease. Meanwhile, a continuous increase in the range of the carbonaceous/carbonate-like species bands was observed. This change of the intensity was interpreted as Boudouard reaction (2 CO → CO₂ + C) on the active sites of Ni.¹⁰⁰

The disproportionation of CO to carbon would result in the partial encapsulation of some Ni sites and a lower intensity for CO_{ad} species. Meanwhile, the produced CO₂ can be adsorbed and formed the carbonate species on the adjacent ZrO₂. Note that one cannot exclude a direct CO adsorption contributing to the formation of carbonaceous/carbonate-like species as described for Ni/MgO. In the desorption process, except for the bands of ionic Niⁿ⁺-CO species, almost all the bands are stable and still present during purging in He (Fig. S5, ESI[†]). Only the linear Ni⁰-CO species lost more than half of its intensity, suggesting the presence of less stable Ni⁰-CO species.

Regarding the Ni/CeO₂ catalyst, CO adsorption led to the formation of the similar species observed for Ni/MgO: linear

Ni⁰-CO (2079 and 2055 cm⁻¹) and bridged (Ni⁰)_x-CO (1979, 1957, 1917 and 1896 cm⁻¹) species.⁹⁶ Various carbonate-like species bands were observed and were ascribed to the CeO₂ support in the frequency region lower than 1800 cm⁻¹. In the CO adsorption process, similarly to the Ni/ZrO₂ sample, a change of the CO_{ad} species intensity was observed. This behaviour becomes even more remarkable on Ni/CeO₂ than Ni/ZrO₂, with a major intensity reduction that drastically reduces the intensity of all CO_{ad} species. Such behaviour could be related to the stronger reducibility of CeO₂ than ZrO₂, which was also proved by the H₂-TPR results (Fig. 4). Compared to Ni/ZrO₂, more oxygen vacancies of Ni/CeO₂ formed during reduction pre-treatment (proved by H₂-TPR) drove a more pronounced charge transfer between Ni and reduced CeO₂ (EMSI effect), resulting in a higher electron density on Ni sites in Ni/CeO₂ catalyst, which leads to a more active CO disproportionation (Fig. 12d). Such concept that charge transfer from oxygen vacancies to adjacent metal sites through EMSI effect has been also demonstrated on Ru/ZrO₂ catalysts both in experimental and theoretical aspects.⁷² Moreover, the formation of carbon layer on the top of Ni⁰ sites during disproportionation would inhibit the further adsorption of CO molecules. Such hypothesis is consistent with the observation that the intensity of CO_{ad} species were much lower than the carbonate bands on Ni/CeO₂ catalyst (Fig. 12c). During the desorption process, almost all CO_{ad} species disappeared (Fig. S6, ESI[†]), suggesting a very labile nature of these species and influenced by a stronger electron density transferred from the reduced CeO₂



support. At the same time, the carbonaceous/carbonate-like species slightly increased in intensity, indicating that some CO molecules could have been re-adsorbed to form carbonaceous/carbonate-like species after being desorbed from Ni⁰ sites.

4. Discussion

The support effect on the Ni,Fe bimetallic catalysts as well as Ni monometallic catalysts in CO₂ hydrogenation (methanation and RWGS) studied here can be summarized and interpreted as follows:

1. On a series of Ni-based (Ni,Fe) catalysts supported on non-reducible MgO, the optimum molar metal composition Ni:Fe for CO₂ methanation was established to be 5:1. In accordance with previous reports on the promoting role of Fe for methanation, an increase in activity and CH₄ selectivity was observed, however, the selectivity around 10% at 350 °C was still low and the activity was not stable. An increased reaction order of H₂ and CO₂ in methanation with the addition of Fe further confirmed the promoting effect, which was not observed for the RWGS reaction.

2. The optimal ratio of Ni:Fe = 5:1 was further introduced to bimetallic Ni,Fe catalysts supported on CeO₂ and ZrO₂ and compared to monometallic counterparts to investigate the combined role of promoter and support effects for these two reducible oxides. CeO₂ was the most active support for both methanation and RWGS reactions, followed by ZrO₂, while MgO-supported bimetallic catalyst showed the lowest catalytic performance. This order corresponds to the reducibility of the supports suggesting metal-support interaction as origin of this effect. On these reducible supports, the presence of Fe even resulted in a bit lower methanation activity while the RWGS reaction was slightly promoted on both CeO₂ and ZrO₂ supported Ni, Fe catalysts. Therefore, different from the non-reducible MgO-supported catalysts, no promoting effect of Fe for methanation was observed on reducible oxides. This can be explained following the explanation of the promoting effect by Grundwaldt *et al.*, who proposed that the redox cycle of Fe at the surface of the Ni particles helps the CO₂ activation. In case of reducible oxides such as ZrO₂ or CeO₂ such activation of CO₂ can happen on oxygen vacancies of the support, and thus no promotion on the metal surface might be required. It was shown previously that oxygen vacancies and metal-zirconia interfacial sites can promote the activation of CO₂.^{5,34,38} Such synergistic interplay of reduced oxide sites for CO₂ activation and metal sites for H₂ activation can also explain the superior stability of the catalysts supported on reducible oxides as no or less carbon species or oxygen-containing coupled products are present on the metal surface leading to deactivation by coking or oxidation.

3. ZrO₂-supported catalysts exhibited the highest CH₄ selectivity (94.4% on Ni/ZrO₂ at 350 °C). Instead of deactivation, they also showed a long activation period (~10 h) during reaction, consistent with the slow formation of oxygen vacancies due to the moderate reducibility. Also, the kinetic parameters were standing out from the other catalysts with

regard to a very high activation energy and a strongly negative reaction order of H₂ for RWGS. In the context of our interpretation, this suggests that CO₂ competes with H₂ for the adsorption sites on the metallic Ni surface. Due to the hydrogen-rich feed stoichiometry, RWGS is inhibited by hydrogen and only little carbonaceous intermediates might be present on the Ni particles. The preferred CO₂ activation on the reduced support sites on the other hand does not lead to CO pathway on the metal sites likely proceeds *via* formate pathway on support sites with adjacent oxygen vacancies to form methane.^{41,78}

4. The *in situ* DRIFTS results demonstrate that MgO, ZrO₂ and CeO₂ have a strong influence on Ni active sites. More specifically, there is a clear difference in the CO_{ad} species formed. The intensity ratio between carbonaceous/carbonates-like species and CO_{ad} bands on each sample after steady-state CO adsorption increased with increasing of the support reducibility (Fig. 12d), which can be interpreted as the Boudouard reaction is more promoted by the presence of oxygen vacancies. Such disproportionation will result carbon layer covered on Ni and prevent the formation of CO_{ad}. The other disproportionation product CO₂ will then be adsorbed on the support, which is consistent with our kinetic analysis. All these findings highlight MSI in reducible oxide-supported Ni-based catalysts, which can be addressed for the difference in their catalytic performance.

5. Conclusion

In conclusion, this study has shown that the well-established promotion effect of Fe on Ni-based catalysts with similar metal loading and metal particle size is present if non-reducible MgO is used as support, which according to literature report can be traced back to facilitated CO₂ activation involving Fe redox cycles. If reducible supports such as ZrO₂ and CeO₂ are involved, no such effect is evident. *In situ* DRIFTS with CO as probe and H₂-TPR as well as CO₂-TPD have demonstrated that the importance of MSI effect in these two supports, resulting into a new mechanism of CO₂ activation on the supports and rendering a promotion on the metal surface by Fe redundant. This hypothesis is consistent with the kinetic parameters reported here, which are most pronounced on Ni/ZrO₂, *i.e.* a large activation energy and a negative reaction order for H₂ in RWGS, while the pre-factor for methanation was still large. For these reasons, Ni/ZrO₂ reached the highest CH₄ selectivity of 94% at 350 °C, which is accompanied by an interesting activation behavior due to the dynamic nature of the under-lying MSI. The catalysts supported on the most reducible oxide CeO₂ with the richest oxygen vacancies showed the highest CO₂ methanation activity.

Data availability

The data supporting this article have been included as part of the ESI.†



Conflicts of interest

There are no conflicts to declare.

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References

- H. Blanco and A. Faaij, *Renewable Sustainable Energy Rev.*, 2018, **81**, 1049–1086.
- X. Ning, R. Lin, R. O'Shea, D. Wall, C. Deng, B. Wu and J. D. Murphy, *iScience*, 2021, **24**, 102998.
- A. Antenucci and G. Sansavini, *Renewable Sustainable Energy Rev.*, 2019, **100**, 33–43.
- R. Estevez, L. Aguado-Deblas, F. M. Bautista, F. J. López-Tenllado, A. A. Romero and D. Luna, *Catalysts*, 2022, **12**, 1555.
- H. Chen, H. Cui, Y. Lv, P. Liu, F. Hao and W. Xiong, *Fuel*, 2022, **314**, 123035.
- R.-P. Ye, J. Ding, W. Gong, M. D. Argyle, Q. Zhong, Y. Wang, C. K. Russell, Z. Xu, A. G. Russell, Q. Li, M. Fan and Y.-G. Yao, *Nat. Commun.*, 2019, **10**, 5698.
- L. Li, W. Zeng, M. Song, X. Wu, G. Li and C. Hu, *Catalysts*, 2022, **12**, 244.
- N. Diyan Mohd Ridzuan, M. S. Shaharun, M. A. Anawar and I. Ud-Din, *Catalysts*, 2022, **12**, 469.
- M. A. Aziz, A. A. Jalil, N. S. Hassan, M. Bin Bahari, A. H. Hatta, T. A. T. Abdullah, N. W. C. Jusoh, H. D. Setiabudi and R. Saravanan, *Process Saf. Environ. Prot.*, 2024, **186**, 1229–1241.
- S. Chen and A. M. Abdel-Mageed, *Int. J. Hydrogen Energy*, 2023, **48**, 24915–24935.
- E. Spennati, P. Riani and G. Garbarino, *Catal. Today*, 2023, **418**, 114131.
- M. Tawalbeh, R. M. N. Javed, A. Al-Othman and F. Almomani, *Energy Convers. Manage.*, 2023, **279**, 116755.
- J. Sehested, K. E. Larsen, A. L. Kustov, A. M. Frey, T. Johannessen, T. Bligaard, M. P. Andersson, J. K. Nørskov and C. H. Christensen, *Top. Catal.*, 2007, **45**, 9–13.
- Y. Feng, L. Shen, W. Zhang, X. Yuan, M. Zhu and J. Xu, *J. CO₂ Util.*, 2024, **80**, 102683.
- A. I. Tsiotsias, N. D. Charisiou, I. V. Yentekakis and M. A. Goula, *Nanomaterials*, 2021, **11**, 28.
- A. L. Kustov, A. M. Frey, K. E. Larsen, T. Johannessen, J. K. Nørskov and C. H. Christensen, *Appl. Catal., A*, 2007, **320**, 98–104.
- Z. Bian, S. Das, M. H. Wai, P. Hongmanorom and S. Kawi, *ChemPhysChem*, 2017, **18**, 3117–3134.
- T. Pu, J. Chen, W. Tu, J. Xu, Y.-F. Han, I. E. Wachs and M. Zhu, *J. Catal.*, 2022, **413**, 821–828.
- J. Li, Q. Xu, Y. Han, Z. Guo, L. Zhao, K. Cheng, Q. Zhang and Y. Wang, *Sci. China: Chem.*, 2023, **66**, 3518–3524.
- M.-A. Serrer, A. Gaur, J. Jelic, S. Weber, C. Fritsch, A. H. Clark, E. Saraçi, F. Studt and J.-D. Grunwaldt, *Catal. Sci. Technol.*, 2020, **10**, 7542–7554.
- M.-A. Serrer, K. F. Kalz, E. Saraçi, H. Lichtenberg and J.-D. Grunwaldt, *ChemCatChem*, 2019, **11**, 5018–5021.
- B. Mutz, M. Belimov, W. Wang, P. Sprenger, M.-A. Serrer, D. Wang, P. Pfeifer, W. Kleist and J.-D. Grunwaldt, *ACS Catal.*, 2017, **7**, 6802–6814.
- A. B. Shirsath, M. L. Schulte, B. Kreitz, S. Tischer, J.-D. Grunwaldt and O. Deutschmann, *Chem. Eng. J.*, 2023, **469**, 143847.
- H. L. Huynh, J. Zhu, G. Zhang, Y. Shen, W. M. Tucho, Y. Ding and Z. Yu, *J. Catal.*, 2020, **392**, 266–277.
- B. Yan, B. Zhao, S. Kattel, Q. Wu, S. Yao, D. Su and J. G. Chen, *J. Catal.*, 2019, **374**, 60–71.
- Z. Zhang, C. Shen, K. Sun, X. Jia, J. Ye and C.-J. Liu, *J. Mater. Chem. A*, 2022, **10**, 5792–5812.
- W. K. Fan and M. Tahir, *J. Environ. Chem. Eng.*, 2021, **9**, 105460.
- Q. Li, C. Wang, H. Wang, J. Chen, J. Chen and H. Jia, *Angew. Chem.*, 2024, **136**, e202318166.
- C. Vogt, E. Groeneveld, G. Kamsma, M. Nachtegaal, L. Lu, C. J. Kiely, P. H. Berben, F. Meirer and B. M. Weckhuysen, *Nat. Catal.*, 2018, **1**, 127–134.
- X. Zhou, G. A. Price, G. J. Sunley and C. Copéret, *Angew. Chem., Int. Ed.*, 2023, **62**, e202314274.
- W. Gac, W. Zawadzki, M. Rotko, M. Greluk, G. Słowik and G. Kolb, *Catal. Today*, 2020, **357**, 468–482.
- P. Frontera, A. Macario, M. Ferraro and P. Antonucci, *Catalysts*, 2017, **7**, 59.
- L. P. Matte, A. S. Kilian, L. Luza, M. C. M. Alves, J. Morais, D. L. Baptista, J. Dupont and F. Bernardi, *J. Phys. Chem. C*, 2015, **119**, 26459–26470.
- K. Larmier, W.-C. Liao, S. Tada, E. Lam, R. Verel, A. Bansode, A. Urakawa, A. Comas-Vives and C. Copéret, *Angew. Chem., Int. Ed.*, 2017, **56**, 2318–2323.
- R. Ye, L. Ma, X. Hong, T. R. Reina, W. Luo, L. Kang, G. Feng, R. Zhang, M. Fan and R. Zhang, *Angew. Chem.*, 2024, **136**, e202317669.
- Y. Ding, J. Chen, X. Lian, Z. Tian, X. Geng, Y. Wang, Y. Liu, W. Wang, M. Wang and Y. Xiao, *Appl. Catal., B*, 2024, **343**, 123508.
- M. A. Abir, R. E. Phillips, J. Z. M. Harrah and M. Ball, *Catal. Sci. Technol.*, 2024, 4506–4521.
- S. Kumari, A. N. Alexandrova and P. Sautet, *J. Am. Chem. Soc.*, 2023, **145**, 26350–26362.



- 39 S. Chen, A. M. Abdel-Mageed, C. Gauckler, S. E. Olesen, I. Chorkendorff and R. J. Behm, *J. Catal.*, 2019, **373**, 103–115.
- 40 Y. Bian, C. Xu, X. Wen, L. Xu, Y. Cui, S. Wang, C.-E. Wu, J. Qiu, G. Cheng and M. Chen, *Fuel*, 2023, **331**, 125755.
- 41 F. Arena, G. Italiano, K. Barbera, S. Bordiga, G. Bonura, L. Spadaro and F. Frusteri, *Appl. Catal., A*, 2008, **350**, 16–23.
- 42 M. Monai, K. Jenkinson, A. E. M. Melcherts, J. N. Louwen, E. A. Irmak, S. Van Aert, T. Altantzis, C. Vogt, W. van der Stam, T. Duchoň, B. Šmíd, E. Groeneveld, P. Berben, S. Bals and B. M. Weckhuysen, *Science*, 2023, **380**, 644–651.
- 43 C. Wei, H. Ding, Z. Zhang, F. Lin, Y. Xu and W. Pan, *Int. J. Hydrogen Energy*, 2024, **58**, 872–891.
- 44 F. Hu, C. Jin, K. H. Lim, C. Li, G. Song, Bella, T. Wang, R. Ye, Z.-H. Lu, G. Feng, R. Zhang and S. Kawi, *Fuel Process. Technol.*, 2023, **250**, 107873.
- 45 G. S. Dhillon, G. Cao and N. Yi, *Catalysts*, 2023, **13**, 1171.
- 46 W. L. Vrijburg, J. W. A. van Helden, A. Parastaev, E. Groeneveld, E. A. Pidko and E. J. M. Hensen, *Catal. Sci. Technol.*, 2019, **9**, 5001–5010.
- 47 M. Jafarbegloo, A. Tarlani, A. W. Mesbah, J. Muzart and S. Sahebdehfar, *Catal. Lett.*, 2016, **146**, 238–248.
- 48 L. Geng, B. Zheng, X. Wang, W. Zhang, S. Wu, M. Jia, W. Yan and G. Liu, *ChemCatChem*, 2016, **8**, 805–811.
- 49 Y. Wei, D. Luo, L. Yan, C. Ma, Z. Fu, L. Guo, M. Cai, S. Sun and C. Zhang, *Catal. Lett.*, 2022, **152**, 1835–1843.
- 50 C. Mebrahtu, S. Perathoner, G. Giorgianni, S. Chen, G. Centi, F. Krebs, R. Palkovits and S. Abate, *Catal. Sci. Technol.*, 2019, **9**, 4023–4035.
- 51 J. Li, J. Li and Q. Zhu, *Chin. J. Chem. Eng.*, 2018, **26**, 2344–2350.
- 52 D. Weber, T. He, M. Wong, C. Moon, A. Zhang, N. Foley, N. J. Ramer and C. Zhang, *Catalysts*, 2021, **11**, 1447.
- 53 C. Janke, M. S. Duyar, M. Hoskins and R. Farrauto, *Appl. Catal., B*, 2014, **152**, 184–191.
- 54 N. L. Visser, S. J. Turner, J. A. Stewart, B. D. Vandegheuchte, J. E. S. van der Hoeven and P. E. de Jongh, *ACS Nano*, 2023, **17**, 14963–14973.
- 55 P. Munnik, M. E. Z. Velthoen, P. E. de Jongh, K. P. de Jong and C. J. Gommers, *Angew. Chem.*, 2014, **126**, 9647–9651.
- 56 M. Mihaylov, K. Hadjiivanov and H. Knözinger, *Catal. Lett.*, 2001, **76**, 59–63.
- 57 M. Agnelli, H. M. Swaan, C. Marquez-Alvarez, G. A. Martin and C. Mirodatos, *J. Catal.*, 1998, **175**, 117–128.
- 58 N. Liu, J. Wei, J. Xu, Y. Yu, J. Yu, Y. Han, K. Wang, J. I. Orege, Q. Ge and J. Sun, *Appl. Catal., B*, 2023, **328**, 122476.
- 59 L. Li, B. Jiang, D. Tang, Z. Zheng and C. Zhao, *Catalysts*, 2018, **8**, 257.
- 60 W. Yu, Q. Zhou, H. Wang, Y. Liu, W. Chu, R. Cai and W. Yang, *J. Mater. Sci.*, 2020, **55**, 2321–2332.
- 61 X. Zhang, R. You, D. Li, T. Cao and W. Huang, *ACS Appl. Mater. Interfaces*, 2017, **9**, 35897–35907.
- 62 Y. Gao, R. Li, S. Chen, L. Luo, T. Cao and W. Huang, *Phys. Chem. Chem. Phys.*, 2015, **17**, 31862–31871.
- 63 W. Zheng, J. Zhang, Q. Ge, H. Xu and W. Li, *Appl. Catal., B*, 2008, **80**, 98–105.
- 64 T. Takeguchi, S.-N. Furukawa and M. Inoue, *J. Catal.*, 2001, **202**, 14–24.
- 65 C. Li, Y. Shi, Z. Zhang, J. Ni, X. Wang, J. Lin, B. Lin and L. Jiang, *J. Energy Chem.*, 2021, **60**, 403–409.
- 66 K. Chang, H. Zhang, M.-J. Cheng and Q. Lu, *ACS Catal.*, 2020, **10**, 613–631.
- 67 M. Carltonbird, S. Eaimsumang, S. Pongstabodee, S. Boonyuen, S. M. Smith and A. Luengnaruemitchai, *Chem. Eng. J.*, 2018, **344**, 545–555.
- 68 X. Li, X. Liu, J. Hao, L. Li, Y. Gao, Y. Gu, Z. Cao and J. Liu, *ACS Omega*, 2022, **7**, 24646–24655.
- 69 A. Parastaev, V. Muravev, E. Huertas Osta, A. J. F. van Hoof, T. F. Kimpel, N. Kosinov and E. J. M. Hensen, *Nat. Catal.*, 2020, **3**, 526–533.
- 70 L. R. Winter, E. Gomez, B. Yan, S. Yao and J. G. Chen, *Appl. Catal., B*, 2018, **224**, 442–450.
- 71 B. Miao, S. S. K. Ma, X. Wang, H. Su and S. H. Chan, *Catal. Sci. Technol.*, 2016, **6**, 4048–4058.
- 72 S. Chen, A. M. Abdel-Mageed, M. Li, S. Cisneros, J. Bansmann, J. Rabeah, A. Brückner, A. Groß and R. J. Behm, *J. Catal.*, 2021, **400**, 407–420.
- 73 X. Jia, X. Zhang, N. Rui, X. Hu and C.-J. Liu, *Appl. Catal., B*, 2019, **244**, 159–169.
- 74 Z. Cheng, B. J. Sherman and C. S. Lo, *J. Chem. Phys.*, 2013, **138**, 014702.
- 75 K. Lorber, J. Zavašnik, I. Arčon, M. Huš, J. Teržan, B. Likozar and P. Djinović, *ACS Appl. Mater. Interfaces*, 2022, **14**, 31862–31878.
- 76 N. Rui, X. Zhang, F. Zhang, Z. Liu, X. Cao, Z. Xie, R. Zou, S. D. Senanayake, Y. Yang, J. A. Rodriguez and C.-J. Liu, *Appl. Catal., B*, 2021, **282**, 119581.
- 77 T. Zhang, W. Wang, F. Gu, W. Xu, J. Zhang, Z. Li, T. Zhu, G. Xu, Z. Zhong and F. Su, *Appl. Catal., B*, 2022, **312**, 121385.
- 78 K. Li and J. G. Chen, *ACS Catal.*, 2019, **9**, 7840–7861.
- 79 J. I. Di Cosimo, C. R. Apesteguía, M. J. L. Ginés and E. Iglesia, *J. Catal.*, 2000, **190**, 261–275.
- 80 D. Wierzbicki, R. Baran, R. Dębek, M. Motak, T. Grzybek, M. E. Gálvez and P. Da Costa, *Int. J. Hydrogen Energy*, 2017, **42**, 23548–23555.
- 81 P. Summa, M. Gajewska, L. Li, C. Hu, B. Samojeden, M. Motak and P. Da Costa, *J. CO₂ Util.*, 2022, **60**, 101983.
- 82 N. J. A. Rahman, A. Ramli, K. Jumbri and Y. Uemura, *Sci. Rep.*, 2019, **9**, 16223.
- 83 Z. Ni, X. Djitchou, X. Gao, J. Wang, H. Liu and Q. Zhang, *Sci. Rep.*, 2022, **12**, 5344.
- 84 W. Liao, C. Tang, H. Zheng, J. Ding, K. Zhang, H. Wang, J. Lu, W. Huang and Z. Zhang, *J. Catal.*, 2022, **407**, 126–140.
- 85 Y. Wang, H. Ban, Y. Wang, R. Yao, S. Zhao, J. Hu and C. Li, *J. Catal.*, 2024, **430**, 115357.
- 86 S. Lin, Z. Li and M. Li, *Fuel*, 2023, **333**, 126369.



- 87 M. B. Jensen, L. G. M. Pettersson, O. Swang and U. Olsbye, *J. Phys. Chem. B*, 2005, **109**, 16774–16781.
- 88 D. Cornu, H. Guesmi, J.-M. Krafft and H. Lauron-Pernot, *J. Phys. Chem. C*, 2012, **116**, 6645–6654.
- 89 R. Ye, L. Ma, X. Hong, T. R. Reina, W. Luo, L. Kang, G. Feng, R. Zhang, M. Fan, R. Zhang and J. Liu, *Angew. Chem.*, 2024, **136**, e202317669.
- 90 Y. Yang, Z. Chai, X. Qin, Z. Zhang, A. Muhetaer, C. Wang, H. Huang, C. Yang, D. Ma, Q. Li and D. Xu, *Angew. Chem., Int. Ed.*, 2022, **61**, e202200567.
- 91 I. Hussain, G. Tanimu, S. Ahmed, C. U. Aniz, H. Alasiri and K. Alhooshani, *Int. J. Hydrogen Energy*, 2023, **48**, 24663–24696.
- 92 B. M. Tackett, E. Gomez and J. G. Chen, *Nat. Catal.*, 2019, **2**, 381–386.
- 93 A. Beck, D. Kazazis, Y. Ekinci, X. Li, E. A. Müller Gubler, A. Kleibert, M.-G. Willinger, L. Artiglia and J. A. van Bokhoven, *ACS Nano*, 2023, **17**, 1091–1099.
- 94 C. Ren, R. Yang, Y. Li and H. Wang, *Rev. Chem. Intermed.*, 2019, **45**, 3019–3032.
- 95 G. Martra, F. Arena, M. Baricco, S. Coluccia, L. Marchese and A. Parmaliana, *Catal. Today*, 1993, **17**, 449–458.
- 96 K. I. Hadjiivanov and G. N. Vayssilov, in *Advances in Catalysis*, Academic Press, 2002, **47**, pp. 307–511.
- 97 G. Busca and V. Lorenzelli, *Mater. Chem.*, 1982, **7**, 89–126.
- 98 S. Chen, T. Cao, Y. Gao, D. Li, F. Xiong and W. Huang, *J. Phys. Chem. C*, 2016, **120**, 21472–21485.
- 99 S. Chen, L. Luo, Z. Jiang and W. Huang, *ACS Catal.*, 2015, **5**, 1653–1662.
- 100 M. B. Jensen, S. Morandi, F. Prinetto, A. O. Sjøstad, U. Olsbye and G. Ghiotti, *Catal. Today*, 2012, **197**, 38–49.

