



Cite this: *Chem. Commun.*, 2024, 60, 14346

# Surface coverage and reconstruction analyses bridge the correlation between structure and activity for electrocatalysis

Zhongyuan Guo,<sup>a,b</sup> Tianyi Wang,<sup>b</sup> Jiang Xu,<sup>\*a</sup> Ang Cao<sup>id,\*cd</sup> and Hao Li<sup>id,\*b</sup>

Electrocatalysis is key to realizing a sustainable future for our society. However, the complex interface between electrocatalysts and electrolytes presents an ongoing challenge in electrocatalysis, hindering the accurate identification of effective/authentic structure–activity relationships and determination of favourable reaction mechanisms. Surface coverage and reconstruction analyses of electrocatalysts are important to address each conjecture and/or conflicting viewpoint on surface-active phases and their corresponding electrocatalytic origin, *i.e.*, so-called structure–activity relationships. In this review, we emphasize the importance of surface states in electrocatalysis experimentally and theoretically, providing guidelines for research practices in discovering promising electrocatalysts. Then, we summarize some recent progress of how surface states determine the adsorption strengths and reaction mechanisms of occurring electrocatalytic reactions, exemplified in the electrochemical oxygen evolution reaction, oxygen reduction reaction, nitrogen reduction reaction, CO<sub>2</sub> reduction reaction, CO<sub>2</sub> and N<sub>2</sub> co-reductions, and hydrogen evolution reaction. Finally, the review proposes deep insights into the *in situ* study of surface states, their efficient building and the application of surface Pourbaix diagrams. This review will accelerate the development of electrocatalysts and electrocatalysis theory by arousing broad consensus on the significance of surface states.

Received 31st July 2024,  
Accepted 27th September 2024

DOI: 10.1039/d4cc03875d

rsc.li/chemcomm

## Introduction

Facing the global energy crises and climate changes, our society is witnessing a desirable revolution of energy infrastructures. Electrocatalysis has great potential to bring us a bright and sustainable future with the target of using renewable resources to convert low-cost feedstocks into value-added chemicals.<sup>1–11</sup> Electrochemical water splitting can produce green hydrogen and oxygen.<sup>12–14</sup> Meanwhile, electrochemical oxygen reduction can convert hydrogen into water without any carbon emissions and simultaneously release electric energy.<sup>11,15–17</sup> The electrochemical nitrogen reduction can catalyze gaseous nitrogen into ammonia under ambient conditions using green electricity without the harsh reaction conditions of the Haber–Boash method.<sup>18–22</sup> The mitigation of unprecedented atmospheric CO<sub>2</sub> levels can also be achieved through electrocatalytic CO<sub>2</sub>

reduction into carbon-containing fuels.<sup>23–25</sup> Furthermore, the coupling of electrocatalytic N<sub>2</sub> and CO<sub>2</sub> co-reduction can be employed to produce urea.<sup>26</sup> Therefore, in past decades, tremendous efforts have been made toward the discovery of promising electrocatalysts, as the core of electrochemical systems, and the development of electrocatalysis theory to guide the rational design of electrocatalyst materials.

The industrial application of electrocatalysis requires electrocatalytic materials with high performance. Currently, various strategies have been developed to design highly efficient and stable electrocatalysts, including defect engineering,<sup>27–29</sup> alloy engineering,<sup>30,31</sup> strain engineering,<sup>32</sup> heterostructure engineering,<sup>33,34</sup> *etc.* These methods have promoted the activity and selectivity of electrocatalysts, although many catalysts cannot reach the required industrial levels. However, their structure–activity relationships remain obscure because surface-active phases with meticulous design are usually different from the resting surface states (*i.e.*, electrochemistry-induced surface coverages and/or the subsequent reconstruction) of catalysts under *in situ* conditions.<sup>35</sup> Due to the electrochemically-driven water activation and ion deposition/dissolution at catalytic interfaces, newly-evolved surface states will emerge with the pre-coverage of some species,<sup>36–43</sup> even resulting from surface reconstructions induced by existing defects (*e.g.*, oxygen and

<sup>a</sup> College of Environmental and Resource Sciences, Zhejiang University, Hangzhou, 310058, China. E-mail: xujiang6@zju.edu.cn, zhongyuanguo2022@163.com

<sup>b</sup> WPI-Advanced Institute for Materials Research (AIMR), Tohoku University, Sendai, 980-8577, Japan. E-mail: li.hao.b8@tohoku.ac.jp, tianyi.wang.c5@tohoku.ac.jp

<sup>c</sup> State Key Laboratory for Clean Energy Utilization, Institute for Thermal Power Engineering, Zhejiang University, Hangzhou 310027, China. E-mail: angc@zju.edu.cn

<sup>d</sup> Inner Mongolia Daqingshan Laboratory, Hohhot 017000, China



phosphorous vacancies).<sup>44,45</sup> Moreover, the water environment can oxidize catalytic materials (like MXenes) to form new species over surfaces.<sup>46</sup> Such complex phenomena complicate the understanding of the origin of electrocatalysis and then limit the advance of electrocatalysts.

In very recent years, the surface state analysis of electrocatalysts has begun attracting attention in the field of electrocatalysis.<sup>36,39,43,44,47–53</sup> In experiments, researchers have found different surface coverages and newly-formed surface-active phases under electrocatalysis compared to their as-designed surfaces.<sup>54–58</sup> These new surface states are highly correlated to their delivered performance. Although some advanced *in situ* and post-reaction characterizations can provide helpful information on surface species and electronic valence of active sites,<sup>59–61</sup> uncovering surface coverages and precise atomic structures/active sites is challenging. Fortunately, the surface Pourbaix diagrams, first proposed in electrocatalysts in the year 2008,<sup>62</sup> have demonstrated the capacity to identify the most energetically-favourable surface states of electrocatalysts under electrocatalytic conditions, which is really helpful to the building of authentic structure–activity relationships. This will be discussed later in detail in this review work. Additionally, the presence of extensive anionic deficiency on surfaces will lead to surface reconstructions, forming new surface-active phases. Based on these scenarios, very recently, we proposed a standard research paradigm for structure–activity relationships in the investigation of electrocatalytic CO<sub>2</sub> reduction over SnO<sub>2</sub>, which is also applicable to other electrocatalytic reactions.<sup>44</sup>

In this review, we will initially discuss how to find stable catalyst surfaces, *i.e.*, surface states, and emphasize the core position of surface state analysis in building structure–activity relationships for electrocatalysis systems. In experiments, the *in situ* characterizations may still provide limited information on the understanding of surface-active species at atomic levels but can act as an important supplement for density functional

theory (DFT)-based theoretical studies. Then we will discuss how surface states determine the adsorption of reaction intermediates and reaction mechanisms rather than as-synthesized surfaces through some electrochemical reactions. Finally, we propose deep insights into the *in situ* study of surface states, the fast and effective building of surface Pourbaix diagrams, and other applications of them.

## The basic research framework for electrocatalysis

Currently, numerous electrocatalysts have been experimentally tested or theoretically screened for accelerating the energy conversion efficiency and reaction kinetics. However, authentic structure–activity relationships existing in different electrocatalytic systems remain obscure although many exciting catalysts have been claimed or predicted. The primary challenge lies in the unknown resting surface of electrocatalysts, stemming from the insufficient consideration of surface state analysis. Our group emphasized the necessity of analysing surface states through different electrochemical reactions with various classic materials, including dual-atom catalysts<sup>37</sup> and metal oxides/nitrides/sulfides.<sup>36,39,44,48,49</sup> In the meantime, García-Melchor and Ciotti have also discussed the importance of surface coverages of electrocatalysts under reaction conditions and that unphysical models may decelerate the discovery process of more effective and cheaper electrocatalysts.<sup>43</sup> To provide a guideline/framework to uncover structure–activity relationships in the community of electrocatalysis, here the key research ingredients for density functional theory (DFT)-based and experimental studies are emphasized as shown in Fig. 1. As for the DFT-based studies, the surface Pourbaix diagrams should emerge as a fundamental tool to illustrate electrochemically-induced surface coverages; then sometimes, some surface species could induce the surface reconstruction

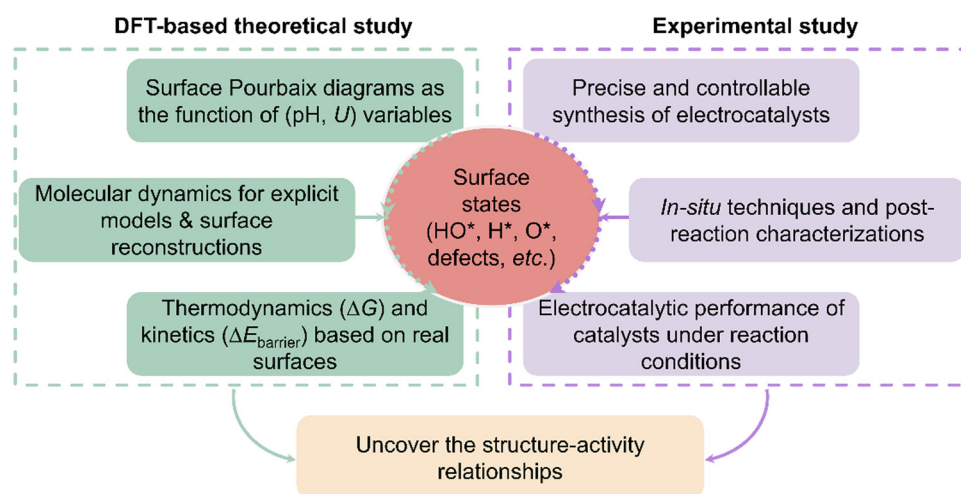


Fig. 1 The proposed basic research framework for theoretical and experimental investigations to build an authentic connection between surface-active structures under the reaction conditions and the delivered performance.



to form a new phase, thus molecular dynamics (MD) simulations using various algorithms, like the constrained genetic algorithm,<sup>63,64</sup> are needed to obtain the most stable surface phase and the evolution process of active sites. Furthermore, the MD simulation of explicit interfacial models is an effective scheme to study the surface states.<sup>46,65,66</sup> Certainly, the surface reconstructions of electrocatalysts can be affected by many factors, including fluctuating electrolyte pH, temperature, applied potentials, and ions. Therefore, advancing more precise MD methodologies that incorporate the aforementioned factors is highly sought after, yet remains a formidable challenge. Following surface state analysis, the analysis of electrocatalytic mechanisms, including thermodynamics and kinetics, can then be conducted based on the real/resting surface states under relevant reaction conditions. In experiments, to illustrate the structure–activity relationship, it is necessary to precisely synthesize the electrocatalysts with controllable surface species and compositions, and coordination environments. Nevertheless, the surface state of active sites during electrochemical reactions may be more complex than typically assumed or designed, further complicating the identification of active sites. Therefore, during the electrocatalytic process, *in situ* advanced characterization techniques should be employed to understand the changes in electronic valences and surface compositions,<sup>67</sup> which determine the activity and selectivity of electrocatalysis. Although *in situ* characterization techniques have been widely employed in various electrocatalytic systems, theoretical DFT-based investigation can provide powerful support to dictate the surface-active phase and the atomic-level understanding of reaction mechanisms. In elucidating structure–activity correlations, the integration of theoretical and experimental approaches is essential to advance the design of efficient electrocatalysts and foster the progression of electrocatalysis theory. In the following section, we will emphasize how surface state analysis aids in decreasing the discrepancy in different electrocatalytic systems, particularly boosting the understanding of surface-active structures and their corresponding performance, including the electrochemical oxygen evolution reaction (OER), oxygen reduction reaction (ORR), nitrogen reduction reaction (NRR), CO<sub>2</sub> reduction reaction (CO<sub>2</sub>RR), CO<sub>2</sub> and N<sub>2</sub> co-reductions, and hydrogen evolution reaction (HER).

## Surface state analysis narrows down the discrepancy in electrocatalysis

### OER electrocatalysis

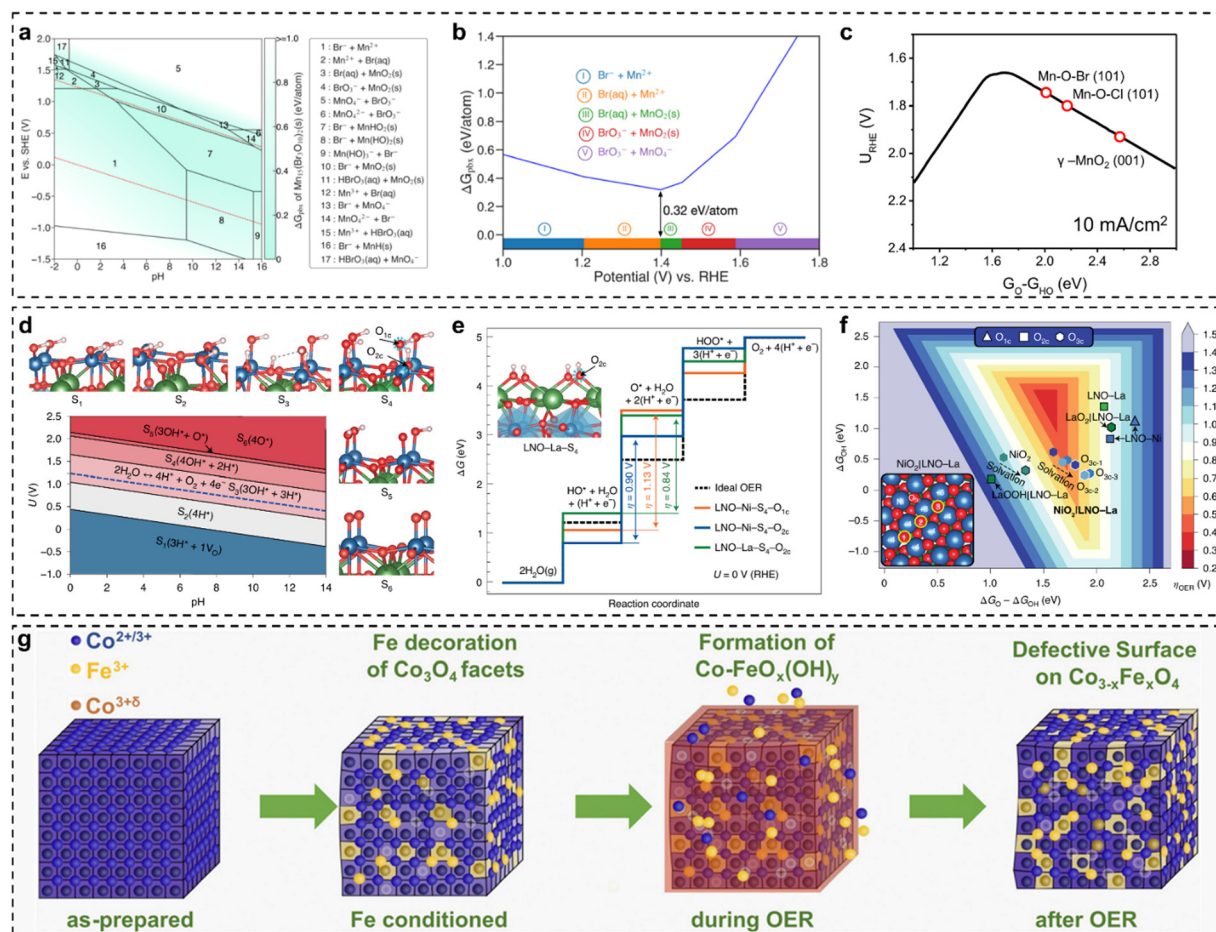
The design of low-cost and effective electrocatalysts for water electrolysis is an open challenge. Many studies have suggested that the OER usually suffers from sluggish kinetics and instability,<sup>13,68–72</sup> significantly making the practical implementation of electrochemical water splitting difficult and inefficient. Very recently, our group employed a doping strategy to design a series of electrocatalysts based on modified zirconium dioxide (ZrO<sub>2</sub>) with excellent thermal stability, identifying that single-atom Fe and Rh dopants showed remarkable activity enhancement towards the OER.<sup>73</sup> The study provides crucial

insights into the design and development of stable, high-activity, and low-cost OER electrocatalysts for water splitting. However, the industrial-scale application of OER electrocatalysts still faces the bottleneck of easily-changed structure–activity relationships, because the interaction between catalysts and electrolyte media is very complicated at large current densities. Achieving long-term stability and high catalytic activity requires a comprehensive understanding of electrochemically-induced surface-active structures and their impact on performance. For example, with knowledge of the surface-catalytic structure–activity relationship, Gu *et al.*<sup>74</sup> tuned the surface structure of FeNi nanoparticles by rational thermal oxidation and/or reductive approaches for the OER, finding that the surface structure of catalysts in a fully oxidized or metallic state is not active, whereas the mixed-phase FeNi alloys demonstrate superior catalytic performance due to the optimized valence states of the Fe/Ni species. Bhattacharyya *et al.*<sup>75</sup> employed the finite cluster models combined with constant potential-DFT and implicit solvation methods to investigate the impact of Fe substitution in cobalt(oxy)hydroxide for the OER. The surface states of the cluster catalyst models reveal that cobalt(oxy)hydroxide predominantly carries a positive charge under alkaline conditions, with nearly complete coverage of hydroxyl ligands coordinated to the Co center. This surface configuration represents the active catalytic phase for the OER. Additionally, the constant potential approach shows that at lower potentials, the O–O bond preferentially forms through the coupling of adjacent Co–OH sites, consistent with experimental findings. Substituting a surface Co atom with Fe significantly enhances the ease of oxidation compared to the unmodified structure. The existence of Fe could reduce the H<sup>+</sup> affinity of the surface-bound –OH moieties, *i.e.*, increased Fe–OH acidity, consequently decreasing the rate-determining step of the deprotonation of the –OH group. Analysis of surface states in finite cluster models confirms experimental results and offers deeper insights into the structure–activity relationship for the OER.

Pan *et al.*<sup>76</sup> synthesized a highly efficient, stable noble-metal-free electrocatalyst to address the limitations in OER activity and stability. The as-synthesized Mn<sub>7.5</sub>O<sub>10</sub>Br<sub>3</sub> catalysts can achieve a low overpotential of ~295 mV at 10 mA cm<sup>–2</sup> and meanwhile maintain good stability operating for > 500 h. DFT-calculated bulk Pourbaix diagrams reveal that the activity and stability of Mn<sub>7.5</sub>O<sub>10</sub>Br<sub>3</sub> arise from the formation of a MnO<sub>x</sub> passivation layer on the surface under reaction conditions (Fig. 2(a)). Compared to Mn<sub>7.5</sub>O<sub>10</sub>Br<sub>3</sub>, Mn<sub>8</sub>O<sub>10</sub>Cl<sub>3</sub> requires a higher operational potential, resulting in a greater thermodynamic driving force to decompose. This explains why Mn–O–Br materials possess superior stability over Mn–O–Cl materials (Fig. 2(b)). The developed microkinetic volcano model predicts that the Mn–O–Br material exhibits the highest OER activity as a function of  $G_{\text{O}} - G_{\text{HO}}$  (Fig. 2(c)). This study reveals that the surface-active phase of Mn<sub>7.5</sub>O<sub>10</sub>Br<sub>3</sub> enhances both stability and OER activity.

Baeumer *et al.*<sup>54</sup> demonstrated how electrochemically driven surface transformations influence the identification of surface-active phases and structure–activity relationships in (001)-terminated LaNiO<sub>3</sub> (LNO) epitaxial thin films, aiding the





**Fig. 2** (a) Calculated Mn–O–Br Pourbaix diagrams with an ion concentration of  $10^{-4}$  M at 298.15 K (note: the Mn ion concentration is based on the ICP-OES result, and the Lake blue colour indicates the stability of Mn<sub>7.5</sub>O<sub>10</sub>Br<sub>3</sub>); (b) calculated Pourbaix decomposition free-energy ( $\Delta G_{\text{pbx}}$ ) of Mn<sub>7.5</sub>O<sub>10</sub>Br<sub>3</sub> and the projection of  $\Delta G_{\text{pbx}}$  onto the potential axis suggests the stable species at corresponding potential regions; (c) the OER activity volcano diagram at  $10 \text{ mA cm}^{-2}$  as the function of  $G_{\text{O}} - G_{\text{H}_2\text{O}}$ ; (d) Surface Pourbaix diagram of LNO–Ni (labels from S<sub>1</sub> to S<sub>6</sub> designate surface coverages from more reduced to more oxidized, V<sub>O</sub> < H < OH < O, V<sub>O</sub> is an oxygen vacancy); (e) calculated OER free-energy diagrams for LNO–Ni–S<sub>4</sub> at the top site (orange) and at the bridge site (blue), and for LNO–La–S<sub>4</sub> at the bridge site (green), are compared with an ideal catalyst (dashed line) at no bias; (f) 2D OER activity volcano map of overpotentials as the function of  $\Delta G_{\text{O}} - \Delta G_{\text{OH}}$  and  $\Delta G_{\text{OH}}$ ; (g) Schematic diagram of the Fe interaction with cubic Co<sub>3</sub>O<sub>4</sub> nanoparticles during the OER.<sup>56</sup>

development of targeted OER electrocatalysts. The study showed that the Ni termination of the as-prepared state was more active than the La termination, with an overpotential difference of  $\sim 150$  mV. The surface Pourbaix diagram of the bulk LNO–Ni, as shown in Fig. 2(d), indicated that as the voltage increased, the LNO–Ni surface transitioned from reduced to more oxidized coverages. Under OER-relevant biases, the Ni-terminated surface was predominantly covered by 1 ML H\* on singly coordinated O<sub>1C</sub> sites and 1/2 ML H\* on doubly coordinated O<sub>2C</sub> sites. Notably, a voltage exceeding 2 V is required to achieve 1 ML O\*<sub>top</sub> (LNO–Ni–S<sub>6</sub>). Based on surface state results, the authors assessed the theoretical OER activity of LNO with various terminations, as illustrated in Fig. 2(e). Using the stable S<sub>4</sub> coverages for both LNO–Ni and LNO–La, it was determined that the HO\*  $\rightarrow$  O\* transition is the potential-limiting step on both surfaces. Specifically, on LNO–Ni, the thermodynamic overpotential for the OER was  $\eta = 1.13$  V at O<sub>1C</sub> and 0.90 V at O<sub>2C</sub>, while on LNO–La, there was only O<sub>2C</sub> with

0.84 V. Although LNO–La was slightly more active than LNO–Ni, both perovskite-type LaO and NiO bulk LNO(001) exhibited unfavourable O\* intermediates, driving the metal sites into unfavourable Ni<sup>4+</sup> and La<sup>3.5+</sup> oxidation states, which contribute to the high overpotentials. Shortly, LNO(001) terminations under OER conditions tend to be highly covered by HO\* and achieve O\* at higher overpotentials. Theoretical OER overpotential volcano mapping (Fig. 2(f)) shows that bulk LNO(001) terminations, including LNO–Ni–S<sub>4</sub> and LNO–La–S<sub>4</sub>, exhibit weak O\* adsorption. However, transitioning to NiO<sub>2</sub>/LNO–La from the as-prepared LNO–Ni surfaces significantly reduces the overpotentials. The layered NiO<sub>2</sub> overlayers, with their triple-coordinated O<sub>3C</sub> sites, stabilize O\* and HO\*, resulting in a shift to a lower and more centered position in the 2D volcano diagram. Finally, it is concluded that electrochemically driven Ni oxyhydroxide-like surfaces enhance electrochemical activity, with surface state analysis confirming the stability of the NiO<sub>2</sub>/LNO–La surface under OER conditions. Moreover, Haase *et al.*<sup>56</sup>



uncovered the link between surface-active structures and the enhancement of Fe decoration over  $\text{Co}_3\text{O}_4$  nano-catalysts. By leveraging various *operando* spectroscopic techniques, the authors demonstrated that Fe decoration enhanced catalytic activity and facilitated the formation of a near-surface-active state, as shown in Fig. 2(g). Fe decoration led to observable near-surface structural transformations in the  $\text{Co}_3\text{O}_4$  coherence length, indicating the formation of a Co-Fe (oxyhydr)oxide. Furthermore, in the presence of Fe decoration, the extent of Co-O near-surface oxidation was less pronounced. Incorporating Fe into a near-surface Co-Fe (oxyhydr)oxide increased the active site density, resulting in enhanced apparent OER activity. The authors also discovered that the Fe ions in the  $\text{Co}_{3-x}\text{Fe}_x\text{O}_4$  near-surface experienced a redox transition and accumulated oxidative charges during the OER. Following the OER, Fe incorporation into Fe-decorated  $\text{Co}_3\text{O}_4$  nano-catalysts led to the formation of  $\text{Co}_{3-x}\text{Fe}_x\text{O}_4$  with metal defects in the spinel near-surface. This work illustrated how *in situ* methods can elucidate surface state evolution and its impact on electrocatalysis, revealing direct links between surface-active structures and electrochemical activity.

### ORR electrocatalysis

Fuel cell technology has the promise of transforming energy infrastructures and significantly lowering carbon emissions, paving the way for a more sustainable and environmentally responsible energy future.<sup>11,77–83</sup> Nonetheless, the structure-activity relationships become more complex due to the adsorption of reaction intermediates on the altered active sites under electrochemical conditions, resulting in discrepancies between theoretically predicted potentials/activities and experimental outcomes. Although fuel cell technology has come a long way, its reliance on Pt-based materials remains a major barrier to large-scale commercialization.<sup>84–87</sup> At present, the ORR faces a bottleneck due to the high overpotential at cathodic electrodes, which in turn reduces the overall energy efficiency of fuel cell systems. Consequently, the development of cost-effective, stable, and efficient alternatives to Pt-group materials is a critical objective. In 2022, Kelly *et al.*<sup>88</sup> incorporated field effects into the computational hydrogen electrode model to explain ORR activity trends over Pt(111), Au(111), and Au(100). They found minimal electric field influence on Pt(111), where \*OH removal limits the  $4\text{e}^-$  ORR, but a significant field dependence on Au(111) and Au(100), where \*OOH formation is rate-limiting. These findings aligned with experimentally observed pH dependencies and highlighted the potential of considering electric field effects to identify more effective ORR catalysts.

Transition metal oxides (TMOs) face intrinsic limitations in the ORR due to the slow O-O bond-breaking step caused by the large dipole moment of the \*O intermediate.<sup>77</sup> However, Yuan *et al.*<sup>89</sup> demonstrated that low-cost, versatile zirconium nitride (ZrN) catalysts exhibit superior  $4\text{e}^-$  ORR performance, even surpassing Pt in alkaline conditions. Despite ZrN's superior ORR activity, its precise mechanism and performance origin remained unclear. To address this, Liu *et al.*<sup>36</sup> employed surface state analysis, electric field effect simulations, and pH-dependent

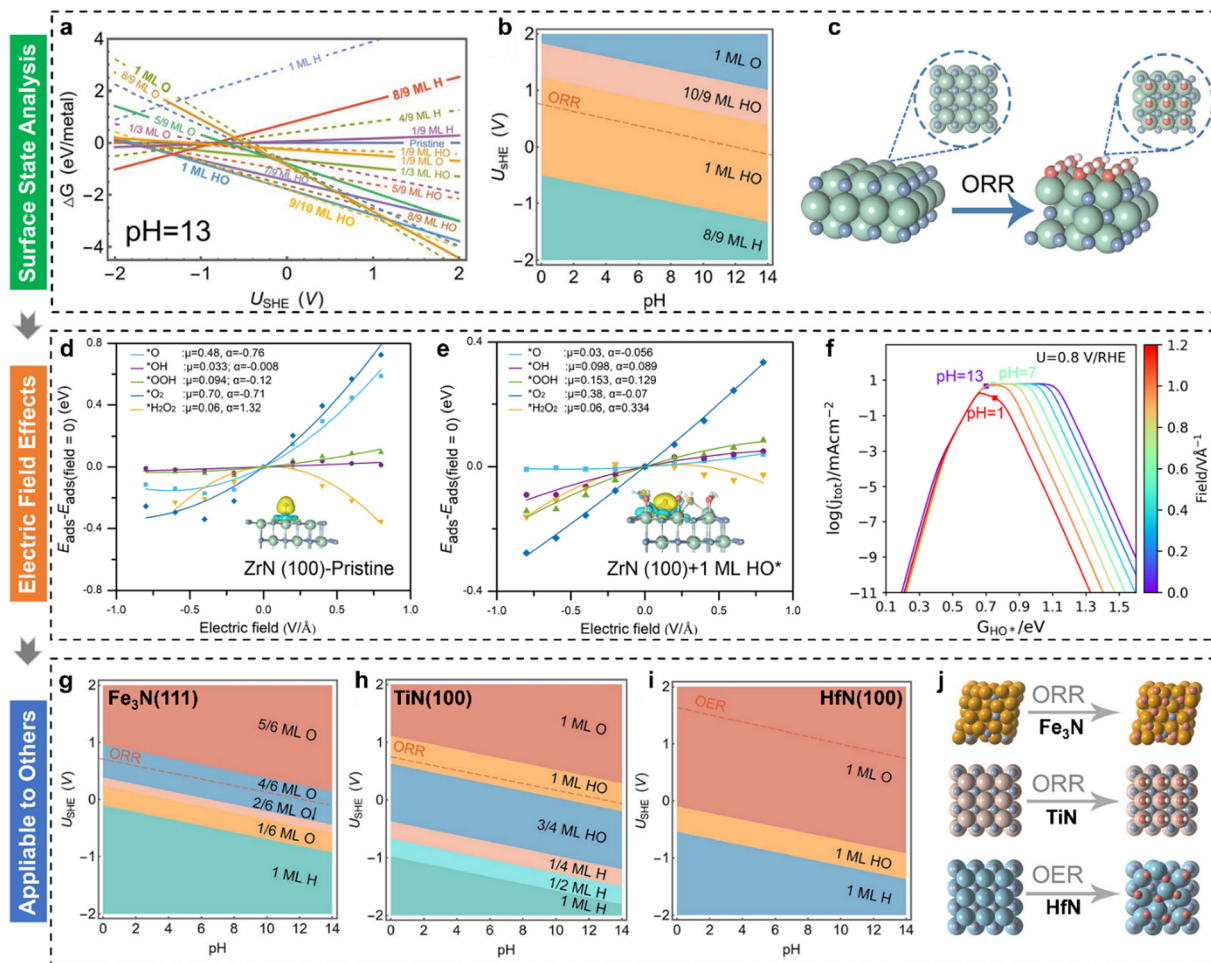
microkinetic modelling to investigate the underlying causes. To explore the origin of ORR performance, the surface Pourbaix diagram for ZrN(100) was calculated (Fig. 3(a)). Under alkaline ORR conditions (pH = 13) and 0.8  $V_{\text{RHE}}$ , ZrN(100) is predicted to be covered by 1 ML  $\text{HO}^*$ . The 2D Pourbaix diagram (Fig. 3(b)) further indicates that this  $\text{HO}^*$  coverage persists across a broad pH range, with the  $\text{HO}^*$ -covered surface shown in Fig. 3(c). Next, the electric field effects on ORR intermediates were examined for both pristine and 1 ML  $\text{HO}^*$ -covered ZrN(100) surfaces (Fig. 3(d) and (e)). It was found that the intrinsic dipole moment ( $\mu$ ) and polarizability ( $\alpha$ ) of  $\text{O}^*$  are significantly reduced on the 1 ML  $\text{HO}^*$ -covered ZrN(100) surface (Fig. 3(e)). The dipole properties of  $\text{O}^*$  are crucial for ORR activity; an ideal catalyst should exhibit minimal dipole moment and polarizability. A larger dipole moment for  $\text{O}^*$  leads to weaker  $\text{O}^*$  bonding, which hinders O-O bond activation and slows the ORR.<sup>77</sup> The pH-dependent volcano activity model (Fig. 3(f)) demonstrates that the 1 ML  $\text{HO}^*$ -covered ZrN(100) exhibits exceptional ORR activity, aligning with experimental observations. The analysis result indicates the importance of determining surface states beforehand. Furthermore, the surface transformation under electrochemical conditions was extended to the  $\text{Fe}_3\text{N}$ , TiN, and HfN in that these three transition metal nitrides had proven to be active in the ORR/OER. Their corresponding 2D surface Pourbaix diagrams are displayed in Fig. 3(g)–(i). All three surfaces underwent electrochemical oxidation under relevant conditions, with their atomic surface structures depicted in Fig. 3(j). These transformed surface-active phases are responsible for the activity in alkaline media, which is consistent with the experimental results.

### NRR electrocatalysis

The nitrogen reduction reaction (NRR) under ambient conditions is viewed as a promising alternative to the industrial Haber-Bosch process, enabling the conversion of inert  $\text{N}_2$  into valuable  $\text{NH}_3$ .<sup>90–94</sup> A major challenge of the NRR lies in the scarcity of highly efficient catalysts for facilitating  $\text{N}_2$  fixation and subsequent hydrogenation. Currently, a wide range of electrocatalysts have been investigated to enhance  $\text{N}_2$  adsorption and activation, including transition metal-based materials,<sup>20,95</sup> p-block-element-based materials,<sup>96–101</sup> single-atom materials,<sup>102–108</sup> dual-atom materials,<sup>109–114</sup> and cluster metal catalysts.<sup>115,116</sup> However, due to the lack of in-depth understanding of surface states, the authentic origin of the electrochemical NRR activity remains obscure. We have suggested that the surface states of dual-atom catalysts usually deviated from the pristine surface under electrocatalytic conditions,<sup>37</sup> and that dual-atom catalysts may be covered by some species at the NRR window.<sup>114</sup> Surface-induced modifications by adsorbed species can reshape the geometric configurations and alter the electronic structures of active sites, leading to variations in reaction mechanisms.

MBenes, as analogs of MXenes, have drawn extensive attention in electrocatalysis. Their surface stability under electrochemical conditions poses a significant challenge, as they are prone to oxidation. Varying surface terminations can markedly affect surface properties and reaction activity. Using  $\text{Zr}_2\text{B}_2$  and





**Fig. 3** (a) Calculated 1D surface Pourbaix diagram as a function of applied potential ( $U_{SHE}$ ) ( $T = 298.15$  K,  $pH = 13$ ); (b) calculated 2D surface Pourbaix diagram as a function of  $U_{SHE}$  and pH ( $T = 298.15$  K); (c) optimized ZrN(100) surface structures before and after the coverage of  $^*\text{OH}$  under the ORR; electric field effects on ORR intermediates with fitted parameters for (d) pristine and (e) 1 ML  $\text{HO}^*$ -covered ZrN(100); (f) derived pH-dependent volcano activity model for the ORR as a function of  $\text{HO}^*$  adsorption; calculated 2D surface Pourbaix diagrams as a function of  $U_{SHE}$  and pH ( $T = 298.15$  K) for (g)  $\text{Fe}_3\text{N}(111)$ , (h)  $\text{TiN}(100)$ , and (i)  $\text{HfN}(100)$ , respectively; (j) corresponding atomic structure evolutions under ORR conditions.<sup>56</sup>

$\text{Hf}_2\text{B}_2$  as examples, Gao *et al.*<sup>117</sup> examined their surface states under typical NRR conditions, revealing that both MBenes become fully terminated by oxygen groups. Additionally, as the pH of the electrolyte increases, a more negative  $U_{SHE}$  potential is required to prevent the oxidation of the bare  $\text{Zr}_2\text{B}_2$  and  $\text{Hf}_2\text{B}_2$  surfaces. Furthermore, various single-atom metals were doped into the oxidized surfaces containing Zr(Hf) vacancies. Among these,  $\text{h-Zr}_2\text{B}_2\text{O}_2\text{-Cr}$  exhibited the highest NRR activity, with the  $^*\text{NH}_2 \rightarrow ^*\text{NH}_3$  transition identified as the potential-determining step, characterized by a maximum  $\Delta G$  of 0.10 eV. However, this study overlooks the surface states of the newly formed doped surfaces, adding uncertainty to the performance of these electrocatalysts under actual reaction conditions.

Recently, our group explored the underlying mechanisms behind the exceptional performance of less-precious  $\text{FeS}_2$  catalysts for the NRR, uncovering the widespread occurrence of *in situ* sulfur vacancy formation in transition metal disulfides ( $\text{TMS}_2$ ) during electrochemical processes.<sup>39</sup> First, according to a

summary of typical  $\text{TMS}_2$  NRR catalysts reported in the literature (Fig. 4(a)), it can be seen that these  $\text{TMS}_2$  preferentially achieve the highest FE or experience a rapid FE growth at a relatively low potential (*i.e.*,  $-0.4$  V<sub>RHE</sub>). To gain deeper insights into the enhanced NRR performance and guide the design of modified  $\text{TMS}_2$  materials, our group selected the  $\text{FeS}_2(111)$  surface—a frequently reported facet known for its higher thermodynamic stability—as a model system to interpret the experimental observations in NRR studies. Fig. 4(b) indicates that the  $\text{FeS}_2$  surface decorated with a single S vacancy is the most stable configuration within the NRR-favored potential range (*e.g.*,  $\sim 0.5$  V<sub>RHE</sub>), suggesting that the S-vacancy can be generated *in situ* on the catalytic surface under these conditions. The formation of S vacancies alters the coordination environment of surface-active sites, leading to distinct NRR catalytic behaviours. Consequently, the entire NRR process on both pristine and single S-vacancy decorated  $\text{FeS}_2$  surfaces has been modelled, as illustrated in Fig. 4(c). It is seen that the S-vacancy has the excellent ability to adsorb  $\text{N}_2$ , but the release

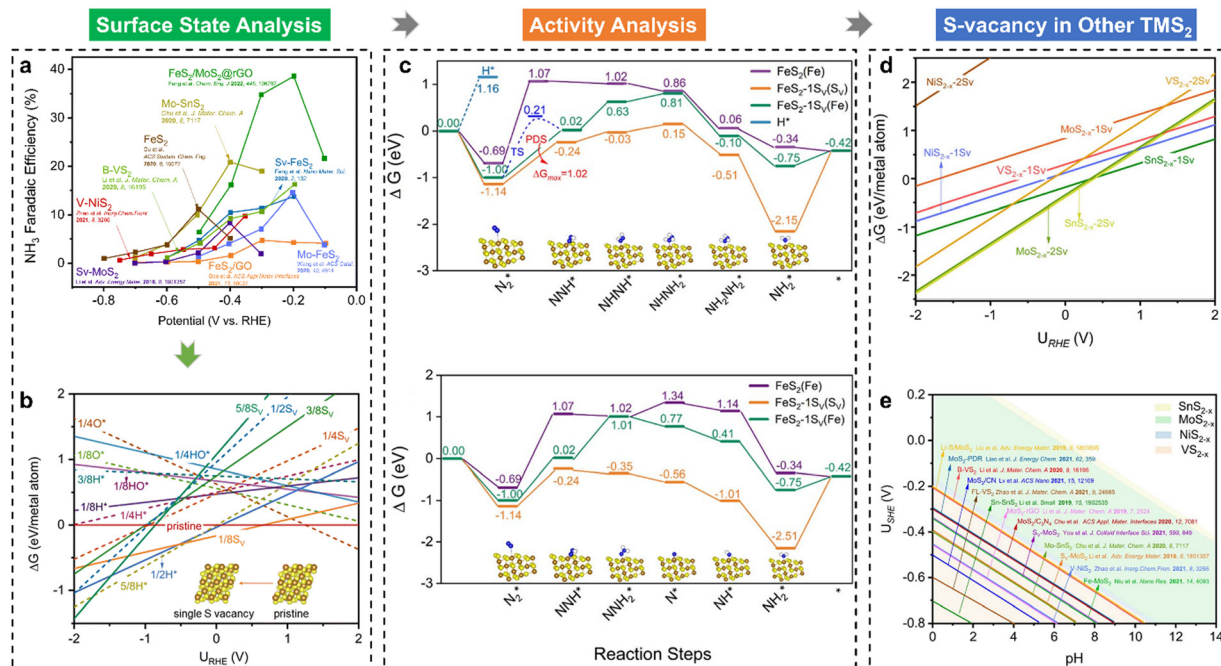


Fig. 4 (a) A summary of the NRR faradaic efficiencies (FEs) of TMS<sub>2</sub> under varying applied potentials; (b) 1D surface Pourbaix diagrams of FeS<sub>2</sub>(111) with the consideration of S<sub>v</sub>, O\*, H\*, and HO\* groups; (c) Gibbs free energy evolutions of the NRR over FeS<sub>2</sub>(111) and FeS<sub>2-x</sub>(111)-1S<sub>v</sub> surfaces, including the alternating and distal pathways; (d) 1D surface Pourbaix diagrams of different TMS<sub>2</sub> surfaces; (e) 2D surface Pourbaix diagram of the S<sub>v</sub> formation as a function of pH and potential, and the reported experimental operating potentials at the highest NH<sub>3</sub> FEs are displayed in the diagram.<sup>39</sup>

of the second NH<sub>3</sub> molecule is significantly endothermic ( $\Delta G = 1.73$  eV). The Fe near the S-vacancy site can also activate N<sub>2</sub> and catalyze the NRR as an active site with the \*N<sub>2</sub> → \*NNH step as the potential-determining step ( $\Delta G = 1.02$  eV). Compared with the stoichiometric pristine FeS<sub>2</sub>(111), the *in situ* generation of an S-vacancy largely boosts the NRR activity and is stable for continuous NH<sub>3</sub> production. In addition to the FeS<sub>2</sub>(111) studied, other TMS<sub>2</sub> materials, including MoS<sub>2</sub>(100), NiS<sub>2</sub>(210), VS<sub>2</sub>(001), SnS<sub>2</sub>(111), and VS<sub>2</sub>(011), were also investigated for the surface state analysis (Fig. 4(d)). As shown, the *in situ* formation of S vacancies on TMS<sub>2</sub> surfaces is energetically favourable under NRR conditions. Moreover, Fig. 4(e) demonstrates the operating potentials of reported TMS<sub>2</sub>-based materials at their highest NRR FEs, which are all located in our predicted potential windows of the S-vacancy generation. Thus, the *in situ* generation of S vacancies is responsible for the observed NRR performance of TMS<sub>2</sub> catalysts.

### CO<sub>2</sub>RR electrocatalysis

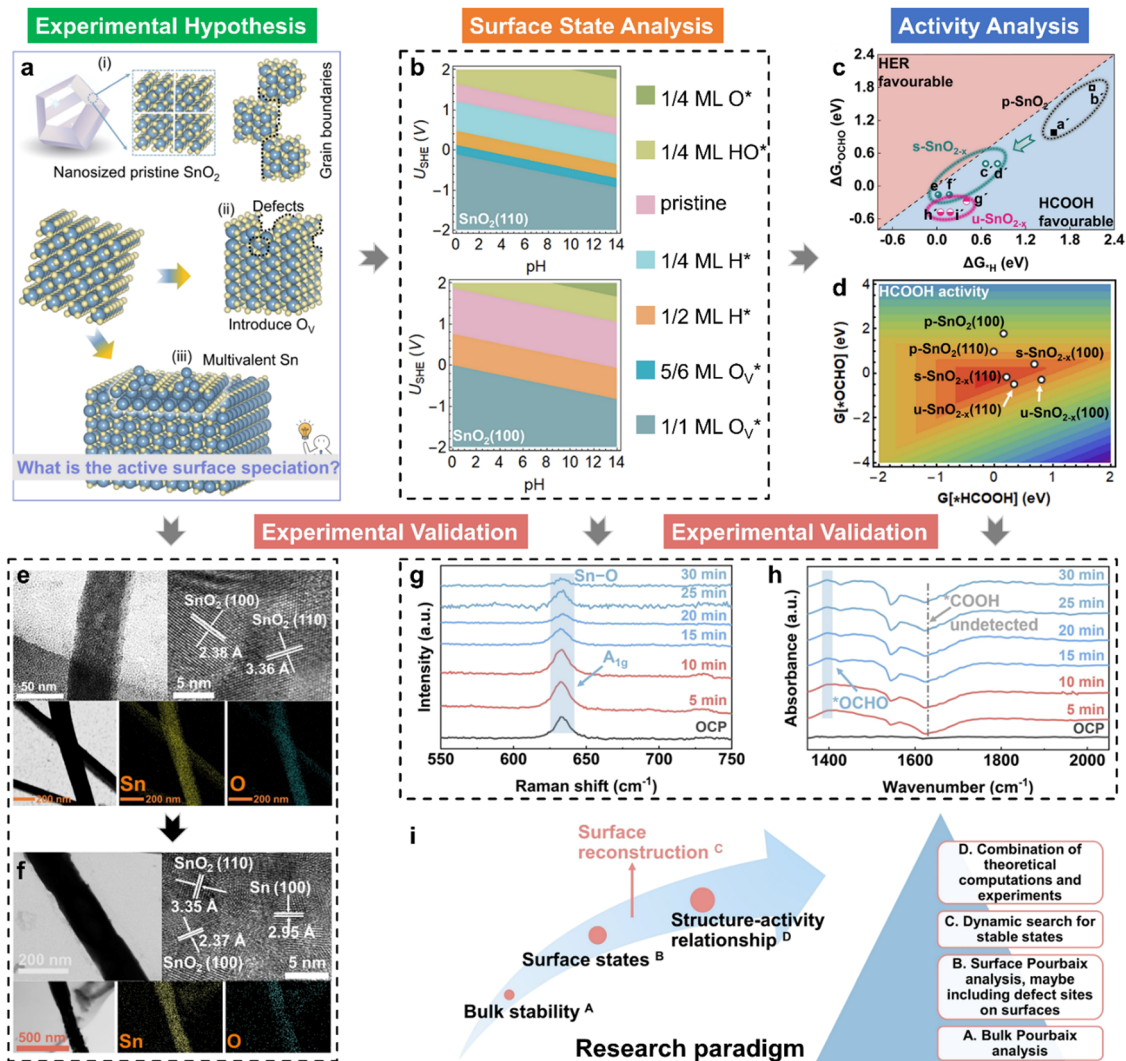
The depletion of fossil fuels has intensified global climate change by driving atmospheric CO<sub>2</sub> levels to unprecedented heights, while their non-renewable nature further underscores the urgent need for a sustainable energy supply to support human development. The electrochemical CO<sub>2</sub> reduction reaction (CO<sub>2</sub>RR) is regarded as an environmentally sustainable and promising strategy to both mitigate excessive CO<sub>2</sub> emissions and convert CO<sub>2</sub> into value-added chemical feedstocks, particularly when integrated with renewable energy sources.<sup>118–124</sup> However, the development of CO<sub>2</sub>RR electrocatalysts with tuneable

activity and selectivity is hindered by the unclear structure–activity relationship that characterizes the CO<sub>2</sub>RR process. For example, (i) why can the SnO<sub>2</sub>-based electrocatalysts deliver a high faradaic efficiency of HCOOH rather than CO,<sup>125–132</sup> and/or what surface species are accountable for the C<sub>1</sub> product distribution?<sup>133–139</sup> (ii) Why does C–C coupling over dual-atom catalysts underperform in CO<sub>2</sub>RR experiments, despite conventional beliefs that dual-atom sites should preferentially facilitate the formation of multi-carbon products?<sup>140–144</sup> These critical challenges drive the need for analyzing surface-active phases under relevant conditions or employing *in situ* surface-sensitive characterization techniques, which can reveal the true origins of the experimental observations.

To understand the CO<sub>2</sub>RR over SnO<sub>2</sub>, as one of the Sn-based materials, we proposed a standard research paradigm to uncover the surface-active species and explore their corresponding CO<sub>2</sub>RR performance.<sup>44</sup> The proposed practice has aided in identifying the resting surface states of SnO<sub>2</sub> under the conditions of the CO<sub>2</sub>RR, corroborated by experimental results, and building the structure–activity relationships for HCOOH and CO production. Firstly, we summarized the current origin of the CO<sub>2</sub>RR over SnO<sub>2</sub>-based materials for the CO<sub>2</sub>RR to HCOOH (Fig. 5(a)), including the (i) interfacial effect of grain boundaries, (ii) the existing oxygen vacancy, and (iii) the multivalent Sn. However, the resting/more realistic surface states of SnO<sub>2</sub> electrocatalysts are unclear, which is closely linked to the electrochemical performance. Therefore, the surface states of SnO<sub>2</sub>(110) and (100), commonly exposed surfaces, were analyzed by the surface Pourbaix diagrams as shown in Fig. 5(b). According to the electrochemical conditions







**Fig. 5** (a) Proposed origin of active  $\text{CO}_2\text{RR}$  over  $\text{SnO}_2$  into  $\text{HCOOH}$ , including the interface effects brought by the grain boundaries of (i) nanosized  $\text{SnO}_2$ , (ii) the existing defects (e.g., O vacancies), and (iii) the multivalent Sn (e.g., Sn/SnO/SnO<sub>2</sub> heterostructures); (b) calculated 2D surface Pourbaix diagrams of  $\text{SnO}_2(110)$  and  $(100)$  with different stable surface states labelled under relevant pH and  $U_{\text{SHE}}$ ; (c) HCOOH selectivity analysis with the comparison of  $\Delta G_{\text{H}^+}$  and  $\Delta G_{\text{OCHO}^*}$ ; (d) the HCOOH activity volcano model for  $\text{SnO}_2$  with different surface states; HRTEM images of the (e) as-synthesized  $\text{SnO}_2$  and (f) "SnO<sub>2</sub>-after" samples, and further the energy dispersive spectroscopy (EDS) elemental mapping of Sn and O after the  $\text{CO}_2\text{RR}$ ; collected *in situ* (g) Raman and (h) ATR-SEIRAS spectra during the  $\text{CO}_2\text{RR}$  on  $\text{SnO}_2$  with a time slot of 5 min; (i) the standard research roadmap to uncover the structure-activity relationships in electrocatalysis.<sup>44</sup>

where the highest faradaic efficiency was achieved, it was found that there would be lots of oxygen vacancies ( $\text{O}_\text{V}^*$ ) over  $\text{SnO}_2(110)$  and  $(100)$  with the coverage of 1 ML  $\text{O}_\text{V}^*$ . The high density of  $\text{O}_\text{V}^*$  then would induce the surface reconstruction to form a new surface-active phase (from  $\text{u-SnO}_{2-x}$  to  $\text{s-SnO}_{2-x}$ ). In Fig. 5(c), the selectivity analysis suggests that the electrochemically induced  $\text{O}_\text{V}^*$  could change the adsorption free energies of both  $\text{H}^*$  and  $\text{OCHO}^*$  ( $\Delta G_{\text{H}^+}$  vs.  $\Delta G_{\text{OCHO}^*}$ ), but still favour the  $\text{HCOOH}$  formation. Fig. 5(d) displays the  $\text{HCOOH}$  activity volcano mapping, where  $\text{s-SnO}_{2-x}(110)$  is the closest to the volcano peak with the highest activity, indicating that the reconstructed surface-active phase, induced by  $\text{O}_\text{V}^*$ , enhances the  $\text{CO}_2\text{RR}$  to  $\text{HCOOH}$ , while pristine  $\text{SnO}_2$  ( $\text{p-SnO}_2$ ) possess the lowest  $\text{HCOOH}$  activity. Furthermore, the  $\text{CO}_2\text{RR}$  experiments

were performed to validate the theoretically uncovered surface species and reaction mechanisms. Fig. 5(e)–(f) display the high-resolution transmission electron microscope (HRTEM) images of the as-synthesized  $\text{SnO}_2$  catalysts before and after the  $\text{CO}_2\text{RR}$ , evidently indicating that metallic Sn was formed due to the generation of an O-vacancy. *In situ* Raman (Fig. 5(g)) and attenuated total reflectance-surface enhanced infrared absorption spectroscopy (ATR-SEIRAS) spectra (Fig. 5(h)) directly suggested that the as-synthesized  $\text{SnO}_2$  catalysts were reduced and the important  $\text{*OCHO}$  intermediate was generated during the  $\text{CO}_2\text{RR}$ , respectively. Finally, a standard research paradigm was proposed by us to uncover the structure-activity relationships in electrocatalysis as shown in Fig. 5(i). To conclude, the study not only uncovers the surface-active species of  $\text{SnO}_2$  but



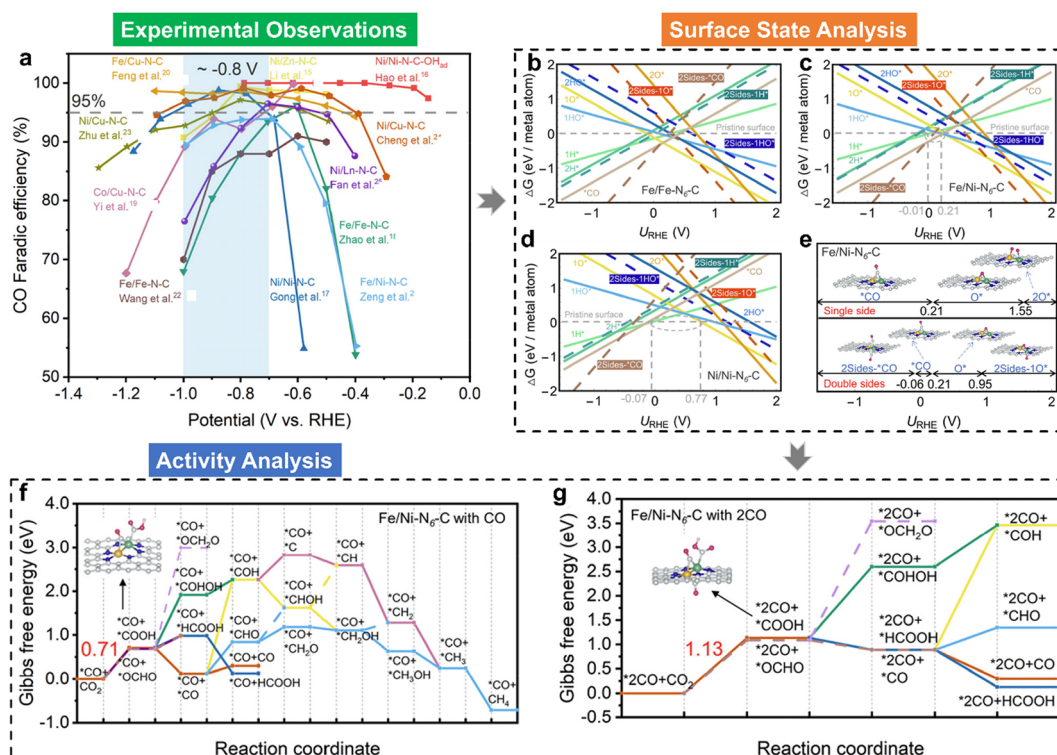
also provides a standard research paradigm for other electrocatalytic systems. The analysis of surface states and surface reconstructions is indispensable, especially in electrocatalysis.

Electrocatalytic C–C coupling represents a crucial advancement in achieving the synthesis of value-added multi-carbon products during the CO<sub>2</sub>RR.<sup>145</sup> However, the dual-atom electrocatalysts, commonly considered as effective electrocatalytic materials for C–C coupling with two active sites for \*CO(\*CHO) intermediates, hardly catalyzed CO<sub>2</sub> into multi-carbon products (see Fig. 6(a)). To study the reason for this unconventional phenomenon, we analyzed the CO<sub>2</sub>RR mechanisms over typical homonuclear (Fe/Fe–N<sub>6</sub>–C and Ni/Ni–N<sub>6</sub>–C) and heteronuclear (Fe/Ni–N<sub>6</sub>–C) dual-atom catalysts initialized by surface state analysis.<sup>48</sup> Fig. 6(b)–(d) illustrate their surface Pourbaix diagrams, where it can be seen that all the dual-atom catalysts studied would be covered by pre-adsorbed CO\* over the bridge site under CO<sub>2</sub>RR conditions, in contrast to the unoccupied dual-atom sites. Fig. 6(e) displays the surface states of Fe/Ni–N<sub>6</sub>–C at different U<sub>RHE</sub>. Notably, at potentials where U<sub>RHE</sub> is below –0.06 V, pre-adsorbed \*CO can form in either single- or double-sided configurations. These pre-adsorbed \*CO species not only alter the electronic structures of the dual-atom active sites but also modify the underlying CO<sub>2</sub>RR mechanisms, as demonstrated in Fig. 6(f) and (g), in which the pre-adsorbed CO\* in the single-side favoured the CO production and hindered the C–C coupling in both thermodynamics and kinetics

and the double-side occupation could lead to the HCOOH formation. As a result of \*CO pre-adsorption-induced poisoning, the formation of multi-carbon products is unfavorable on these dual-atom catalysts. Our study reveals why C–C coupling remains difficult although the precise dual-atom sites are designed. Meanwhile, the work also indicates that regulating the adsorption strength of CO\* on bridge sites may be an effective way to improve the performance of CO<sub>2</sub>RR over dual-atom catalysts.

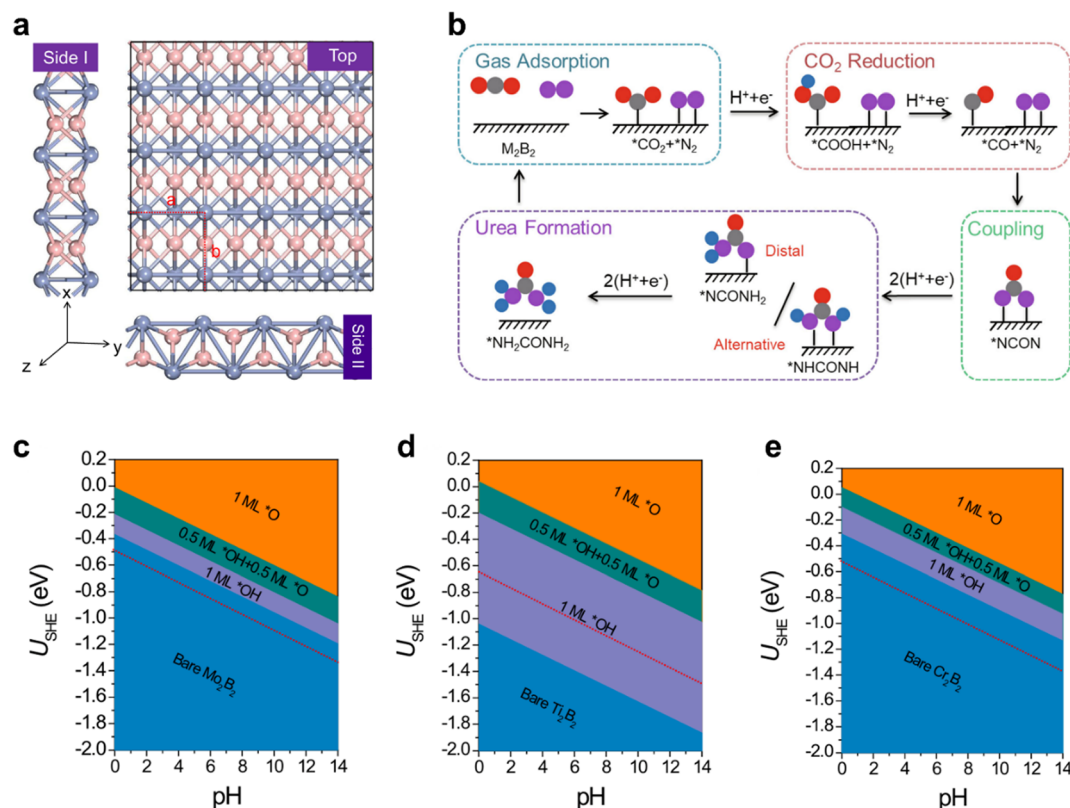
### CO<sub>2</sub> + N<sub>2</sub> electrocatalysis for urea

Urea (CO(NH<sub>2</sub>)<sub>2</sub>), as the first organic compound derived from inorganic raw materials, is an extremely important material in the chemical industry, and is widely used as N<sub>2</sub> fertilizer due to its high nitrogen content (46%). Current industrial urea production relies on extreme reaction conditions, with ammonia predominantly derived from the Haber–Bosch process, which contributes significantly to the energy crisis and exacerbates climate change. Electrochemical urea synthesis *via* CO<sub>2</sub> and N<sub>2</sub> coupling under ambient conditions offers a sustainable and highly efficient approach to urea production.<sup>146–148</sup> However, the large-scale deployment of this reaction is limited by the lack of effective electrocatalysts and complex reaction mechanisms. Zhu *et al.*<sup>26</sup> theoretically investigated the catalytic activity of experimentally available metal borides (MBenes), *i.e.*, Mo<sub>2</sub>B<sub>2</sub>, Ti<sub>2</sub>B<sub>2</sub>, and Cr<sub>2</sub>B<sub>2</sub>, and analyzed their surface structures/states



**Fig. 6** (a) Comparison of the faradaic efficiency of CO (FE<sub>CO</sub>) for various experimentally-reported dual-atom catalysts, where the FE<sub>CO</sub> data as a function of U<sub>RHE</sub> were obtained from their corresponding literatures; calculated 1D surface Pourbaix diagrams of (b) Fe/Fe–N<sub>6</sub>–C, (c) Fe/Ni–N<sub>6</sub>–C, and (d) Ni/Ni–N<sub>6</sub>–C dual-atom catalysts, respectively, where the solid and dashed lines represent single- and double-side adsorption; (e) the surface states of Fe/Ni–N<sub>6</sub>–C at various potential windows, where inserts indicate the atomic configurations; Gibbs free-energy evolutions of the CO<sub>2</sub>RR over Fe/Ni–N<sub>6</sub>–C with pre-covered \*CO at (f) single- and (g) double-sides, respectively.<sup>48</sup>





**Fig. 7** (a) The top and side views of three MBenes with the red dashed lines of a unit cell (pink: boron atoms; blue: metal atoms); (b) the schematic diagram of the reaction mechanism of urea production through the  $\text{*CO}$  and  $\text{*N}_2$  coupling (the gray, red, pink, and blue balls represent C, O, N, and H atoms, respectively); the surface Pourbaix diagrams of (c)  $\text{Mo}_2\text{B}_2$ , (d)  $\text{Ti}_2\text{B}_2$ , and (e)  $\text{Cr}_2\text{B}_2$  with the SHE and pH as variables and the red dashed line represents the limiting potential of urea electro-synthesis; the thermodynamically stable surface states are highlighted by orange ( $\text{*O}$ ), green ( $\text{*O} + \text{*OH}$ ), and purple ( $\text{*OH}$ ).<sup>26</sup>

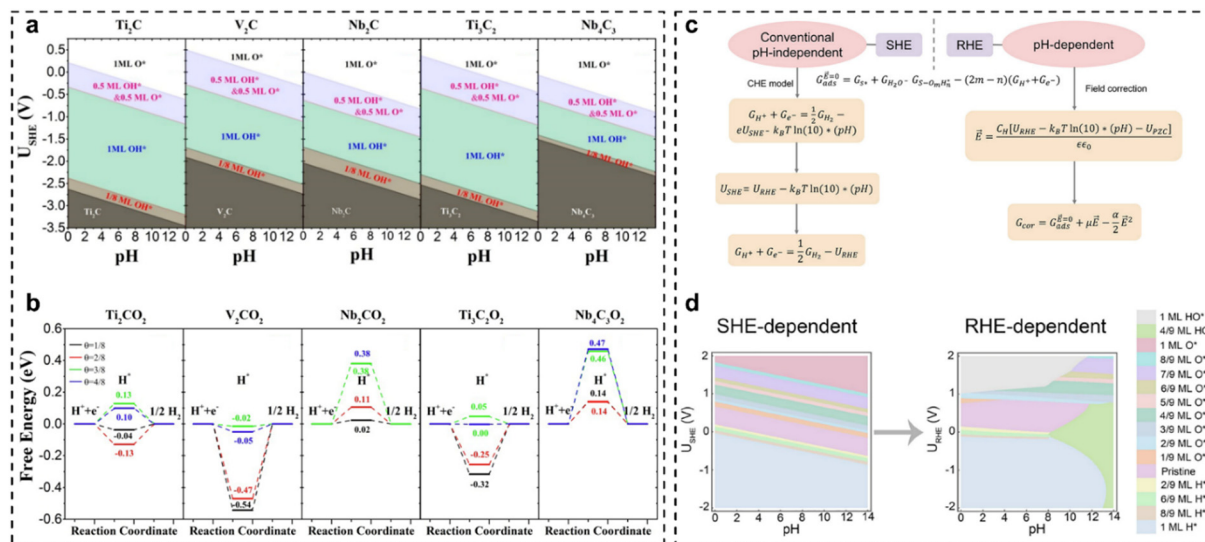
and activity origin for the simultaneous electrocatalytic coupling of  $\text{N}_2$  and  $\text{CO}_2$  to produce urea (Fig. 7). These MBenes catalysts possess metal terminations (Fig. 7(a)), which can act as active sites for the activation of  $\text{*N}_2$  and  $\text{*CO}_2$ . The urea synthesis mechanism is demonstrated in Fig. 7(b), including the key coupling step of  $\text{*CO} + \text{*N}_2 \rightarrow \text{*NCON}$ . As an optimal catalyst candidate, the catalytic surface with the necessary active sites must remain stable and resist coverage by  $\text{*O}/\text{*OH}$  species under operating conditions. Therefore, the surface Pourbaix diagrams of the three MBenes were constructed (Fig. 7(c)–(e)), which revealed that 2D  $\text{Mo}_2\text{B}_2$  and  $\text{Cr}_2\text{B}_2$  possess high electrochemical stability against the surface oxidation, while the 2D  $\text{Ti}_2\text{B}_2$  would be occupied by  $\text{*O}/\text{*OH}$  species under working conditions. Although all three MBenes can effectively activate  $\text{N}_2$  and  $\text{CO}_2$  on their basal surface, the  $\text{*CO}$  intermediate can be easily formed through the  $\text{*CO}_2 \rightarrow \text{*CO}$  process, and the limiting potentials of urea electro-synthesis, comparable to that of Pd–Cu alloys, are in the appropriate range of  $-0.49$  to  $-0.65$  V, the  $\text{Ti}_2\text{B}_2$  catalyst is not an optimal candidate for urea production because the active basal plane will be covered by  $\text{*O}/\text{*OH}$  groups. Therefore, this work strongly reaffirms that surface states under operating conditions must be carefully considered when designing active and selective catalysts. The interaction between the designed surface and the electrolyte

medium can lead to the formation of new, complex phases, resulting in unanticipated structure–activity relationships that may deviate from initial expectations.

### HER electrocatalysis

Developing efficient catalysts for the hydrogen evolution reaction (HER) is significant for a sustainable hydrogen economy.<sup>149–152</sup> MXenes (transition metal carbides, carbonitrides, and nitrides), an important family of two-dimensional (2D) materials, have drawn tremendous attention for HER electrocatalysis due to the large surface area, hydrophilic nature, fast electron transport, and diverse chemical compositions.<sup>150–154</sup> Usually, MXenes are synthesized by etching the A elements in their bulk MAX phases, which leads to the inevitable coverage by O, OH, and/or F atoms.<sup>155–162</sup> Additionally, our report has suggested that the pH and potentials of interest can induce different surface-covered species, and even vacancies and H-covered surfaces at a low potential window.<sup>40</sup> The coverage of different species/newly-formed active sites can significantly impact the surface properties of MXenes and HER performance. Therefore, Gao *et al.*<sup>151</sup> systematically studied the HER activity of MXenes with the termination of O/OH groups. The surface Pourbaix diagrams indicated that  $\text{Ti}_2\text{C}$ ,  $\text{V}_2\text{C}$ , and  $\text{Ti}_3\text{C}_2$  would be terminated by 0.5 ML  $\text{O}^*$  and 0.5 ML  $\text{HO}^*$  groups, while  $\text{Nb}_2\text{C}$  and  $\text{Nb}_4\text{C}_3$  would be fully terminated by  $\text{O}^*$  groups under the





**Fig. 8** (a) Calculated surface Pourbaix diagrams of  $\text{Ti}_2\text{C}$ ,  $\text{V}_2\text{C}$ ,  $\text{Nb}_2\text{C}$ ,  $\text{Ti}_3\text{C}_2$ , and  $\text{Nb}_4\text{C}_3$ ; (b) free-energy evolutions of the HER at different  $\text{H}^*$  coverages ( $\theta = 1/8, 2/8, 3/8$ , and  $4/8$ ) occurring on  $\text{Ti}_2\text{CO}_2$ ,  $\text{V}_2\text{CO}_2$ ,  $\text{Nb}_2\text{CO}_2$ ,  $\text{Ti}_3\text{C}_2\text{O}_2$ , and  $\text{Nb}_4\text{C}_3\text{O}_2$  under standard conditions.<sup>151</sup> (c) Schematic illustration of the RHE scale-dependent method; (d) calculated classic SHE-dependent and RHE-dependent surface Pourbaix diagrams for Pt(111).<sup>50</sup>

reaction potential window, as shown in Fig. 8(a). Then, the hydrogen adsorption free energy ( $\Delta G_{\text{H}^*}$ ) at varying coverages over  $\text{O}^*$ -terminated MXenes was calculated, as shown in Fig. 8(b). Ideally, the  $\Delta G_{\text{H}^*}$  should be close to 0 for a catalyst with high HER performance. In Fig. 8(b), it is found that the average  $\Delta G_{\text{H}^*}$  over the  $\text{O}^*$ -termination of  $\text{Ti}_2\text{C}$  for  $\theta$  ranging from  $1/8$  to  $4/8$  is close to 0, indicating that  $\text{Ti}_2\text{CO}_2$  could offer the optimal binding strength for  $\text{H}^*$  species over a wide range of hydrogen coverages. Therefore,  $\text{Ti}_2\text{CO}_2$  is anticipated to exhibit the highest HER activity among the  $\text{O}^*$ -terminated MXenes studied. The research highlights that  $\text{O}^*$ -termination modulates H adsorption strength, influencing reaction kinetics in the HER.

In catalyst design and electrocatalysis mechanism analysis, the impact of electrochemically induced surface coverage is a crucial factor. The surface Pourbaix diagram is a key tool, offering critical insights into adsorbate coverage from electrochemical potential-driven water activation. However, the classic standard hydrogen electrode (SHE)-dependent surface Pourbaix diagrams are sometimes inconsistent with the experimentally observed surface coverages. This is because some critical factors, illustrated in Fig. 8(c), need to be considered to minimize the discrepancy between theoretical predictions and experimental observations. Within this context, very recently, our group revisited the surface Pourbaix diagram of Pt(111) based on the reversible hydrogen electrode (RHE)-scale by considering the energetics computed by DFT with the Bayesian Error Estimation Functional with van der Waals corrections (BEEF-vdW), the electric field effects, the derived adsorption-induced dipole moment and polarizability, and the potential of zero-charge. As shown in Fig. 8(d), the RHE-dependent surface Pourbaix diagram effectively reduces discrepancies between the theoretical prediction and experimental results for Pt(111),<sup>163</sup> particularly under neutral-alkaline, moderate-potential conditions. To be specific, under alkaline HER conditions at slightly

negative potentials on the RHE scale, the Pt(111) surface is primarily covered by  $\text{HO}^*$  rather than being pristine or  $\text{H}^*$ -covered, which can be helpful to illustrate the alkaline HER mechanism and design optimum catalysts for accelerated HER rate.

## Summary and outlook

In electrocatalysis, surface states of electrocatalysts are complex and often change with the reaction conditions, consequently delivering distinct experimental observations. Thus, effective electrocatalyst design and structure-activity relationship understanding must prioritize surface state analysis, as exemplified in the above-mentioned OER/ORR, NRR,  $\text{CO}_2\text{RR}$ ,  $\text{CO}_2 + \text{N}_2$  electrocatalysis, and HER. As for the surface state analysis, the surface Pourbaix diagram can be analyzed based on the SHE and/or RHE scale. Sometimes, calculating the pH-dependent surface Pourbaix diagram at the RHE scale can minimize the discrepancy between theoretical predictions and experimental observations, particularly under high-pH conditions. In short, the effect of electrochemically-induced surface states should not be overlooked when designing catalysts and identifying electrocatalysis mechanisms.

Additionally, *in situ* characterization techniques would exert positive effects on the understanding of surface structures, because the application of *in situ* characterization techniques can provide real-time insights into the evolution of surface states under operational conditions. For example, Chen *et al.*,<sup>55</sup> using *operando* NAP-XPS, demonstrated that nickel gadolinium-doped ceria cermet electrodes undergo partial reduction, forming a core-shell structure with metallic nickel in the shell and  $\text{NiO}$  beneath. This structure improved *i-V* polarization curves. Thus, beyond theoretical DFT frameworks, advanced *in situ* techniques





have significant potential for revealing surface states, bridging the gap between electrocatalysis performance and catalyst surface structures.

Given the importance of surface structures, developing precise and efficient methods for calculating surface Pourbaix diagrams is highly desirable. Currently, surface Pourbaix diagram analysis is prohibitively expensive for real-scale systems.<sup>164</sup> To address challenges in large catalytic systems, a bond-type embedded crystal graph convolutional neural network (BECGCNN) model was recently developed.<sup>165</sup> With the aid of the BECGCNN model, Bang *et al.*<sup>165</sup> constructed reliable Pourbaix diagrams for large-size nanoparticles (approximately 4.8 nm in diameter, up to 6525 atoms). Additionally, Han *et al.*<sup>166</sup> proposed a Bayesian evolutionary multitasking framework to accelerate the understanding of surface phase diagrams. The integration of Bayesian statistics with evolutionary multitasking enables the efficient generation of surface phase diagrams for complex systems.

Finally, surface state analysis of catalysts has been instrumental in revealing the stability of catalytic materials, greatly aiding in the design of dissolution- and oxidation-resistant catalysts. For example, Li *et al.*<sup>167</sup> discovered that adding Ag atoms to the Pd nanoparticle surface can inhibit surface Pd oxidation and dissolution through surface Pourbaix diagrams. Therefore, surface state analysis under reaction conditions is essential not only for the understanding of the structure–activity relationships, but also for designing optimal catalysts for industrial applications.

## Author contributions

Z. G. drafted the manuscript. Z. G., T. W., J. X., A. C., and H. L. revised the manuscript together, and J. X., A. C., and H. L. supervised the project.

## Data availability

No primary research results, software or code have been included and no new data were generated or analysed as part of this review.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This work is supported by the National Key Research and Development Program of China (No. 2021YFA1202700), the Postdoctoral Fellowship Program of CPSF under the Grant Number (GZC20241472), JSPS KAKENHI (No. JP23K13703 & JP23KF0102), and AIMR Fusion Research.

## References

- V. Rosca, M. Duca, M. T. de Groot and M. T. Koper, *Chem. Rev.*, 2009, **109**, 2209–2244.
- E. J. Popczun, C. G. Read, C. W. Roske, N. S. Lewis and R. E. Schaak, *Angew. Chem., Int. Ed.*, 2014, **53**, 5427–5430.
- J. Zhang and L. Dai, *ACS Catal.*, 2015, **5**, 7244–7253.
- C. Hu and L. Dai, *Angew. Chem., Int. Ed.*, 2016, **55**, 11736–11758.
- Z. W. Seh, J. Kibsgaard, C. F. Dickens, I. Chorkendorff, J. K. Nørskov and T. F. Jaramillo, *Science*, 2017, **355**, eaad4998.
- H. Jin, C. Guo, X. Liu, J. Liu, A. Vasileff, Y. Jiao, Y. Zheng and S.-Z. Qiao, *Chem. Rev.*, 2018, **118**, 6337–6408.
- C. Hu and L. Dai, *Adv. Mater.*, 2019, **31**, 1804672.
- Y. Sun, S. Sun, H. Yang, S. Xi, J. Gracia and Z. J. Xu, *Adv. Mater.*, 2020, **32**, 2003297.
- Y. Li, H. Wang, C. Priest, S. Li, P. Xu and G. Wu, *Adv. Mater.*, 2021, **33**, 2000381.
- X. Jia, Z. Yu, F. Liu, H. Liu, D. Zhang, E. Campos dos Santos, H. Zheng, Y. Hashimoto, Y. Chen, L. Wei and H. Li, *Adv. Sci.*, 2024, **11**, 2305630.
- D. Zhang, Z. Wang, F. Liu, P. Yi, L. Peng, Y. Chen, L. Wei and H. Li, *J. Am. Chem. Soc.*, 2024, **146**, 3210–3219.
- J. Rossmeisl, A. Logadottir and J. K. Nørskov, *Chem. Phys.*, 2005, **319**, 178–184.
- E. Fabbri and T. J. Schmidt, *ACS Catal.*, 2018, **8**, 9765–9774.
- Z. Pei, H. Zhang, D. Luan and X. W. Lou, *Matter*, 2023, **6**, 4128–4144.
- J. K. Nørskov, J. Rossmeisl, A. Logadottir, L. Lindqvist, J. R. Kitchin, T. Bligaard and H. Jónsson, *J. Phys. Chem. B*, 2004, **108**, 17886–17892.
- V. Stamenkovic, B. S. Mun, K. J. J. Mayrhofer, P. N. Ross, N. M. Markovic, J. Rossmeisl, J. Greeley and J. K. Nørskov, *Angew. Chem., Int. Ed.*, 2006, **45**, 2897–2901.
- C. F. Dickens, C. Kirk and J. K. Nørskov, *J. Phys. Chem. C*, 2019, **123**, 18960–18977.
- R. Shi, X. Zhang, G. I. N. Waterhouse, Y. Zhao and T. Zhang, *Adv. Energy Mater.*, 2020, **10**, 2000659.
- W. Tong, B. Huang, P. Wang, Q. Shao and X. Huang, *Natl. Sci. Rev.*, 2021, **8**.
- T. Wang, Z. Guo, X. Zhang, Q. Li, A. Yu, C. Wu and C. Sun, *J. Mater. Sci. Technol.*, 2023, **140**, 121–134.
- A. R. Singh, B. A. Rohr, J. A. Schwalbe, M. Cargnello, K. Chan, T. F. Jaramillo, I. Chorkendorff and J. K. Nørskov, *ACS Catal.*, 2017, **7**, 706–709.
- B. Chang, Z. Cao, Y. Ren, C. Chen, L. Cavallo, F. Raziq, S. Zuo, W. Zhou, Y. Han and H. Zhang, *ACS Nano*, 2024, **18**, 288–298.
- J. T. Feaster, C. Shi, E. R. Cave, T. Hatsukade, D. N. Abram, K. P. Kuhl, C. Hahn, J. K. Nørskov and T. F. Jaramillo, *ACS Catal.*, 2017, **7**, 4822–4827.
- L. D. Chen, M. Urushihara, K. Chan and J. K. Nørskov, *ACS Catal.*, 2016, **6**, 7133–7139.
- A. A. Peterson, F. Abild-Pedersen, F. Studt, J. Rossmeisl and J. K. Nørskov, *Energy Environ. Sci.*, 2010, **3**, 1311–1315.
- X. Zhou, Y. Jing and Y. Li, *Nat. Commun.*, 2021, **12**, 4080.
- A. Zhang, Y. Liang, H. Zhang, Z. Geng and J. Zeng, *Chem. Soc. Rev.*, 2021, **50**, 9817–9844.
- Y. Jia, Y. Zhang, H. Xu, J. Li, M. Gao and X. Yang, *ACS Catal.*, 2024, **14**, 4601–4637.
- X. Yan, L. Zhuang, Z. Zhu and X. Yao, *Nanoscale*, 2021, **13**, 3327–3345.
- P. Zhang, S. Hong, N. Song, Z. Han, F. Ge, G. Dai, H. Dong and C. Li, *Chin. Chem. Lett.*, 2024, **35**, 109073.
- J.-T. Ren, L. Chen, H.-Y. Wang and Z.-Y. Yuan, *Chem. Soc. Rev.*, 2023, **52**, 8319–8373.
- X. Yang, Y. Wang, X. Tong and N. Yang, *Adv. Energy Mater.*, 2022, **12**, 2102261.
- S. Pal, T. Ahmed, S. Khatun and P. Roy, *ACS Appl. Energy Mater.*, 2023, **6**, 7737–7784.
- Z. Zhang, P. Liu, Y. Song, Y. Hou, B. Xu, T. Liao, H. Zhang, J. Guo and Z. Sun, *Adv. Sci.*, 2022, **9**, 2204297.
- C. Lucky and M. Schreier, *ACS Nano*, 2024, **18**, 6008–6015.
- H. Liu, D. Zhang, S. M. Holmes, C. D'Agostino and H. Li, *Chem. Sci.*, 2023, **14**, 9000–9009.
- W. Yang, Z. Jia, B. Zhou, L. Wei, Z. Gao and H. Li, *Commun. Chem.*, 2023, **6**, 6.
- M. Cho, J. T. Song, S. Back, Y. Jung and J. Oh, *ACS Catal.*, 2018, **8**, 1178–1185.



- 39 T. Wang, Z. Guo, H. Oka, A. Kumatani, C. Liu and H. Li, *J. Mater. Chem. A*, 2024, **12**, 8438–8446.
- 40 H. Liu, X. Jia, A. Cao, L. Wei, C. D'agostino and H. Li, *J. Chem. Phys.*, 2023, **158**, 124705.
- 41 J. Duan, S. Chen, C. A. Ortiz-Ledón, M. Jaroniec and S.-Z. Qiao, *Angew. Chem., Int. Ed.*, 2020, **59**, 8181–8186.
- 42 A. S. Dobrota, N. V. Skorodumova, S. V. Mentus and I. A. Pašti, *Electrochim. Acta*, 2022, **412**, 140155.
- 43 A. Ciotti and M. García-Melchor, *Curr. Opin. Electrochem.*, 2023, **42**, 101402.
- 44 Z. Guo, Y. Yu, C. Li, E. Campos dos Santos, T. Wang, H. Li, J. Xu, C. Liu and H. Li, *Angew. Chem., Int. Ed.*, 2024, **63**, e202319913.
- 45 X. Wang, X. Liu, S. Wu, K. Liu, X. Meng, B. Li, J. Lai, L. Wang and S. Feng, *Nano Energy*, 2023, **109**, 108292.
- 46 P. Hou, Y. Tian, Y. Xie, F. Du, G. Chen, A. Vojvodic, J. Wu and X. Meng, *Angew. Chem., Int. Ed.*, 2023, **62**, e202304205.
- 47 M. López, K. S. Exner, F. Viñes and F. Illas, *Adv. Theor. Simul.*, 2023, **6**, 2200217.
- 48 W. Yang, Z. Jia, B. Zhou, L. Chen, X. Ding, L. Jiao, H. Zheng, Z. Gao, Q. Wang and H. Li, *ACS Catal.*, 2023, **13**, 9695–9705.
- 49 J.-T. Jiang, Z. Guo, S.-K. Deng, X. Jia, H. Liu, J. Xu, H. Li and L.-H. Cheng, *ChemSusChem*, 2024, e202400827, DOI: [10.1002/cssc.202400827](https://doi.org/10.1002/cssc.202400827).
- 50 H. Liu, D. Zhang, Y. Wang and H. Li, *Langmuir*, 2024, **40**, 7632–7638.
- 51 S. Pan, H. Li, T. Wang, Y. Fu, S. Wang, Z. Xie, L. Wei, H. Li and N. Li, *ACS Catal.*, 2024, 13814–13824, DOI: [10.1021/acscatal.4c03088](https://doi.org/10.1021/acscatal.4c03088).
- 52 H. Du, T. Wang, M. Li, Z. Yin, R. Lv, M. Zhang, X. Wu, Y. Tang, H. Li and G. Fu, *Adv. Mater.*, 2024, 2411090, DOI: [10.1002/adma.202411090](https://doi.org/10.1002/adma.202411090).
- 53 A. Martini, D. Hursán, J. Timoshenko, M. Rüscher, F. Haase, C. Rettenmaier, E. Ortega, A. Etxebarria and B. Roldan Cuenya, *J. Am. Chem. Soc.*, 2023, **145**, 17351–17366.
- 54 C. Baeumer, J. Li, Q. Lu, A. Y.-L. Liang, L. Jin, H. P. Martins, T. Duchoň, M. Glöck, S. M. Gericke, M. A. Wohlgemuth, M. Giesen, E. E. Penn, R. Dittmann, F. Gunkel, R. Waser, M. Bajdich, S. Nemsák, J. T. Mefford and W. C. Chueh, *Nat. Mater.*, 2021, **20**, 674–682.
- 55 D. Chen, D. K. Niakolas, V. Papaefthimiou, E. Ioannidou, S. G. Neophytides and S. Zafeirotas, *J. Catal.*, 2021, **404**, 518–528.
- 56 F. T. Haase, E. Ortega, S. Saddeler, F.-P. Schmidt, D. Cruz, F. Scholten, M. Rüscher, A. Martini, H. S. Jeon, A. Herzog, U. Hejral, E. M. Davis, J. Timoshenko, A. Knop-Gericke, T. Lunkenbein, S. Schulz, A. Bergmann and B. Roldan Cuenya, *Energy Environ. Sci.*, 2024, **17**, 2046–2058.
- 57 P. P. Lopes, D. Y. Chung, X. Rui, H. Zheng, H. He, P. Farinazzo Bergamo Dias Martins, D. Strmcnik, V. R. Stamenkovic, P. Zapol, J. F. Mitchell, R. F. Klie and N. M. Markovic, *J. Am. Chem. Soc.*, 2021, **143**, 2741–2750.
- 58 Y. Sun, J. Wang, S. Xi, J. Shen, S. Luo, J. Ge, S. Sun, Y. Chen, J. V. Hanna, S. Li, X. Wang and Z. J. Xu, *Nat. Commun.*, 2023, **14**, 2467.
- 59 J.-Y. Ye, Y.-X. Jiang, T. Sheng and S.-G. Sun, *Nano Energy*, 2016, **29**, 414–427.
- 60 Y. Yang, S. Louisia, S. Yu, J. Jin, I. Roh, C. Chen, M. V. Fonseca Guzman, J. Feijóo, P.-C. Chen, H. Wang, C. J. Pollock, X. Huang, Y.-T. Shao, C. Wang, D. A. Muller, H. D. Abruña and P. Yang, *Nature*, 2023, **614**, 262–269.
- 61 Y. Deng, J. Zhao, S. Wang, R. Chen, J. Ding, H.-J. Tsai, W.-J. Zeng, S.-F. Hung, W. Xu, J. Wang, F. Jaouen, X. Li, Y. Huang and B. Liu, *J. Am. Chem. Soc.*, 2023, **145**, 7242–7251.
- 62 H. A. Hansen, J. Rossmeisl and J. K. Nørskov, *Phys. Chem. Chem. Phys.*, 2008, **10**, 3722–3730.
- 63 E. Campos dos Santos, R. Sato, K. Kisu, K. Sau, X. Jia, F. Yang, S.-I. Orimo and H. Li, *Chem. Mater.*, 2023, **35**, 5996–6004.
- 64 Q. Zhu, A. R. Oganov, C. W. Glass and H. T. Stokes, *Acta Cryst.*, 2012, **B68**, 215–226.
- 65 H. Zhao, X. Lv and Y.-G. Wang, *Adv. Sci.*, 2023, **10**, 2303677.
- 66 Z. Zhang, J. Li and Y.-G. Wang, *Acc. Chem. Res.*, 2024, **57**, 198–207.
- 67 X. Cao, D. Tan, B. Wulan, K. S. Hui, K. N. Hui and J. Zhang, *Small Methods*, 2021, **5**, 2100700.
- 68 S. Zhao, Y. Wang, J. Dong, C.-T. He, H. Yin, P. An, K. Zhao, X. Zhang, C. Gao, L. Zhang, J. Lv, J. Wang, J. Zhang, A. M. Khattak, N. A. Khan, Z. Wei, J. Zhang, S. Liu, H. Zhao and Z. Tang, *Nat. Energy*, 2016, **1**, 16184.
- 69 X. Ren, T. Wu, Y. Sun, Y. Li, G. Xian, X. Liu, C. Shen, J. Gracia, H.-J. Gao, H. Yang and Z. J. Xu, *Nat. Commun.*, 2021, **12**, 2608.
- 70 S. Lyu, C. Guo, J. Wang, Z. Li, B. Yang, L. Lei, L. Wang, J. Xiao, T. Zhang and Y. Hou, *Nat. Commun.*, 2022, **13**, 6171.
- 71 Y. Hao, Z. Guo, H. Cheng, C. Yao, S. Cheng, L. Yi and H. Li, *J. Colloid Interface Sci.*, 2023, **650**, 1949–1957.
- 72 C. Feng, M. B. Faheem, J. Fu, Y. Xiao, C. Li and Y. Li, *ACS Catal.*, 2020, **10**, 4019–4047.
- 73 Y. Zhang, T. Wang, L. Mei, R. Yang, W. Guo, H. Li and Z. Zeng, *Nano-Micro Lett.*, 2024, **16**, 180.
- 74 X. Gu, Z. Liu, M. Li, J. Tian and L. Feng, *Appl. Catal., B*, 2021, **297**, 120462.
- 75 K. Bhattacharyya and A. A. Auer, *J. Phys. Chem. C*, 2022, **126**, 18623–18635.
- 76 S. Pan, H. Li, D. Liu, R. Huang, X. Pan, D. Ren, J. Li, M. Shakouri, Q. Zhang, M. Wang, C. Wei, L. Mai, B. Zhang, Y. Zhao, Z. Wang, M. Graetzel and X. Zhang, *Nat. Commun.*, 2022, **13**, 2294.
- 77 H. Li, S. Kelly, D. Guevarra, Z. Wang, Y. Wang, J. A. Haber, M. Anand, G. T. K. K. Gunasooriya, C. S. Abraham, S. Vijay, J. M. Gregoire and J. K. Nørskov, *Nat. Catal.*, 2021, **4**, 463–468.
- 78 S. Yang, Y. J. Tak, J. Kim, A. Soon and H. Lee, *ACS Catal.*, 2017, **7**, 1301–1307.
- 79 A. Kulkarni, S. Siahrostami, A. Patel and J. K. Nørskov, *Chem. Rev.*, 2018, **118**, 2302–2312.
- 80 X. X. Wang, D. A. Cullen, Y.-T. Pan, S. Hwang, M. Wang, Z. Feng, J. Wang, M. H. Engelhard, H. Zhang, Y. He, Y. Shao, D. Su, K. L. More, J. S. Spendelow and G. Wu, *Adv. Mater.*, 2018, **30**, 1706758.
- 81 Y. Wang, Y.-J. Tang and K. Zhou, *J. Am. Chem. Soc.*, 2019, **141**, 14115–14119.
- 82 L. Yang, J. Shui, L. Du, Y. Shao, J. Liu, L. Dai and Z. Hu, *Adv. Mater.*, 2019, **31**, 1804799.
- 83 D. Zhang, Y. Hirai, K. Nakamura, K. Ito, Y. Matsuo, K. Ishibashi, Y. Hashimoto, H. Yabu and H. Li, *Chem. Sci.*, 2024, **15**, 5123–5132.
- 84 R. Lin, X. Cai, H. Zeng and Z. Yu, *Adv. Mater.*, 2018, **30**, 1705332.
- 85 J. Wu and H. Yang, *Acc. Chem. Res.*, 2013, **46**, 1848–1857.
- 86 H. Niu, C. Xia, L. Huang, S. Zaman, T. Maiyalagan, W. Guo, B. You and B. Y. Xia, *Chin. J. Catal.*, 2022, **43**, 1459–1472.
- 87 L. Wang, A. Holewinski and C. Wang, *ACS Catal.*, 2018, **8**, 9388–9398.
- 88 S. R. Kelly, C. Kirk, K. Chan and J. K. Nørskov, *J. Phys. Chem. C*, 2020, **124**, 14581–14591.
- 89 Y. Yuan, J. Wang, S. Adimi, H. Shen, T. Thomas, R. Ma, J. P. Attfield and M. Yang, *Nat. Mater.*, 2020, **19**, 282–286.
- 90 D. Bao, Q. Zhang, F. L. Meng, H. X. Zhong, M. M. Shi, Y. Zhang, J. M. Yan, Q. Jiang and X. B. Zhang, *Adv. Mater.*, 2017, **29**.
- 91 G. F. Chen, S. Ren, L. Zhang, H. Cheng, Y. Luo, K. Zhu, L. X. Ding and H. Wang, *Small Methods*, 2018, **3**, 1800337.
- 92 S. K. Sahoo, J. Heske, M. Antonietti, Q. Qin, M. Oschatz and T. D. Kühne, *ACS Appl. Energy Mater.*, 2020, **3**, 10061–10069.
- 93 H. Jin, S. S. Kim, S. Venkateshalu, J. Lee, K. Lee and K. Jin, *Adv. Sci.*, 2023, **10**, 2300951.
- 94 D. Chandra, R. Xing, T. Xu, Q. Liu, Y. Luo, S. Liu, R. A. Tufa, T. H. Dolla, T. Montini and X. Sun, *Chem. Commun.*, 2021, **57**, 7335–7349.
- 95 D. Yang, T. Chen and Z. Wang, *J. Mater. Chem. A*, 2017, **5**, 18967–18971.
- 96 Z. Chen, J. Zhao, L. Yin and Z. Chen, *J. Mater. Chem. A*, 2019, **7**, 13284–13292.
- 97 Z. Guo, L. Jasin Arachchige, S. Qiu, X. Zhang, Y. Xu, S. J. Langford and C. Sun, *Nanoscale*, 2021, **13**, 14935–14944.
- 98 Z. Guo, T. Wang, H. Liu, S. Qiu, X. Zhang, Y. Xu, S. J. Langford and C. Sun, *Appl. Surf. Sci.*, 2022, **572**, 151441.
- 99 Z. Guo, S. Qiu, H. Li, Y. Xu, S. J. Langford and C. Sun, *Diamond Relat. Mater.*, 2021, **111**, 108210.
- 100 C. Liu, Q. Li, J. Zhang, Y. Jin, D. R. MacFarlane and C. Sun, *J. Phys. Chem. C*, 2018, **122**, 25268–25273.
- 101 C. Liu, Q. Li, C. Wu, J. Zhang, Y. Jin, D. R. MacFarlane and C. Sun, *J. Am. Chem. Soc.*, 2019, **141**, 2884–2888.
- 102 Z. Guo, C. Liu, C. Sun, J. Xu, H. Li and T. Wang, *ChemCatChem*, 2023, **15**, e202300669.
- 103 C. Liu, H. Zheng, T. Wang, Z. Guo, F. Zhu, H. Xie, G. Qin, H. Li and S. Li, *J. Mater. Sci. Technol.*, 2023, **159**, 244–250.
- 104 L. Zhang, H. Zhou, X. Yang, S. Zhang, H. Zhang, X. Yang, X. Su, J. Zhang and Z. Lin, *Angew. Chem., Int. Ed.*, 2023, **135**, e202217473.



- 105 Y. Li, J. Li, J. Huang, J. Chen, Y. Kong, B. Yang, Z. Li, L. Lei, G. Chai, Z. Wen, L. Dai and Y. Hou, *Angew. Chem., Int. Ed.*, 2021, **60**, 9078–9085.
- 106 X.-F. Li, Q.-K. Li, J. Cheng, L. Liu, Q. Yan, Y. Wu, X.-H. Zhang, Z.-Y. Wang, Q. Qiu and Y. Luo, *J. Am. Chem. Soc.*, 2016, **138**, 8706–8709.
- 107 X. Chen, X. Zhao, Z. Kong, W.-J. Ong and N. Li, *J. Mater. Chem. A*, 2018, **6**, 21941–21948.
- 108 C. Liu, Q. Li, J. Zhang, Y. Jin, D. R. MacFarlane and C. Sun, *J. Mater. Chem. A*, 2019, **7**, 4771–4776.
- 109 C. Liu, H. Zheng, T. Wang, X. Zhang, Z. Guo and H. Li, *Phys. Chem. Chem. Phys.*, 2023, **25**, 13126–13135.
- 110 X. Lv, W. Wei, B. Huang, Y. Dai and T. Frauenheim, *Nano Lett.*, 2021, **21**, 1871–1878.
- 111 B. Wang, S. Huang, L. Yang, Q. Fu and Y. Bu, *J. Phys. Chem. C*, 2021, **125**, 14253–14262.
- 112 M. Liang, X. Shao and H. Lee, *Chem. – Eur. J.*, 2024, **30**, e202302843.
- 113 X. Liu, Z. Wang, G. Feng, Y. Sun, X. Zhang, X. Chen, R. Sa, Q. Li, C. Sun and Z. Ma, *Chem. – Eur. J.*, 2024, **30**, e202303148.
- 114 Y. Zhang, Z. Yu, F. She, L. Wei, Z. Zeng and H. Li, *J. Colloid Interface Sci.*, 2023, **640**, 983–989.
- 115 C. Yao, N. Guo, S. Xi, C.-Q. Xu, W. Liu, X. Zhao, J. Li, H. Fang, J. Su, Z. Chen, H. Yan, Z. Qiu, P. Lyu, C. Chen, H. Xu, X. Peng, X. Li, B. Liu, C. Su, S. J. Pennycook, C.-J. Sun, J. Li, C. Zhang, Y. Du and J. Lu, *Nat. Commun.*, 2020, **11**, 4389.
- 116 H. Yang, D. Luo, R. Gao, D. Wang, H. Li, Z. Zhao, M. Feng and Z. Chen, *Phys. Chem. Chem. Phys.*, 2021, **23**, 16707–16717.
- 117 Y. Gao, E. Wang, Y. Zheng, J. Zhou and Z. Sun, *Energy Mater. Adv.*, 2023, **4**, 0039.
- 118 W. Liu, P. Zhai, A. Li, B. Wei, K. Si, Y. Wei, X. Wang, G. Zhu, Q. Chen, X. Gu, R. Zhang, W. Zhou and Y. Gong, *Nat. Commun.*, 2022, **13**, 1877.
- 119 J. Zhang, C. Guo, S. Fang, X. Zhao, L. Li, H. Jiang, Z. Liu, Z. Fan, W. Xu, J. Xiao and M. Zhong, *Nat. Commun.*, 2023, **14**, 1298.
- 120 S. Nitopi, E. Bertheussen, S. B. Scott, X. Liu, A. K. Engstfeld, S. Horch, B. Seger, I. E. L. Stephens, K. Chan, C. Hahn, J. K. Nørskov, T. F. Jaramillo and I. Chorkendorff, *Chem. Rev.*, 2019, **119**, 7610–7672.
- 121 S. Garg, M. Li, A. Z. Weber, L. Ge, L. Li, V. Rudolph, G. Wang and T. E. Rufford, *J. Mater. Chem. A*, 2020, **8**, 1511–1544.
- 122 J.-J. Wang, X.-P. Li, B.-F. Cui, Z. Zhang, X.-F. Hu, J. Ding, Y.-D. Deng, X.-P. Han and W.-B. Hu, *Rare Met.*, 2021, **40**, 3019–3037.
- 123 B. Chang, H. Pang, F. Raziq, S. Wang, K.-W. Huang, J. Ye and H. Zhang, *Energy Environ. Sci.*, 2023, **16**, 4714–4758.
- 124 X. Zhang, W. Huang, L. Yu, M. García-Melchor, D. Wang, L. Zhi and H. Zhang, *Carbon Energy*, 2024, **6**, e362.
- 125 X. Wang, D. Wu, X. Kang, J. Zhang, X.-Z. Fu and J.-L. Luo, *J. Energy Chem.*, 2022, **71**, 159–166.
- 126 B. Ning, M. Liu, Y. Hu, H. Jiang and C. Li, *Dalton Trans.*, 2022, **51**, 3512–3519.
- 127 C. Lin, Z. Xu, X. Kong, H. Zheng, Z. Geng and J. Zeng, *ChemNano-Mat*, 2022, **8**, e202200020.
- 128 M. K. Kim, H. Lee, J. H. Won, W. Sim, S. J. Kang, H. Choi, M. Sharma, H.-S. Oh, S. Ringe, Y. Kwon and H. M. Jeong, *Adv. Funct. Mater.*, 2022, **32**, 2107349.
- 129 Y. Zhang, H. Xu, D. Niu, X. Zhang and Y. Zhang, *ChemSusChem*, 2021, **14**, 2769–2779.
- 130 B. Zhang, S. Chen, B. Wulan and J. Zhang, *Chem. Eng. J.*, 2021, **421**, 130003.
- 131 Y. Yuan, K. Sheng, G. Zhuang, Q. Li, C. Dou, Q.-J. Fang, W.-W. Zhan, H. Gao, D. Sun and X. Han, *Chem. Commun.*, 2021, **57**, 8636–8639.
- 132 J. Li, J. Jiao, H. Zhang, P. Zhu, H. Ma, C. Chen, H. Xiao and Q. Lu, *ACS Sustainable Chem. Eng.*, 2020, **8**, 4975–4982.
- 133 K. V. Daele, D. Arenas-Esteban, D. Choukroun, S. Hoekx, A. Rossen, N. Daems, D. Pant, S. Bals and T. Breugelmans, *ChemElectroChem*, 2023, **10**, e202201024.
- 134 Y. Zang, T. Liu, H. Li, P. Wei, Y. Song, C. Cheng, D. Gao, Y. Song, G. Wang and X. Bao, *Chem. Eng. J.*, 2022, **446**, 137444.
- 135 C. Salvini, M. Re Fiorentin, F. Risplendi, F. Raffone and G. Cicero, *J. Phys. Chem. C*, 2022, **126**, 14441–14447.
- 136 Z. Yang, C. Yang, J. Han, W. Zhao, S. Shao, S. Li, H. Gao, H. Xie and X. Zhang, *J. Mater. Chem. A*, 2021, **9**, 19681–19686.
- 137 X. Cao, B. Wulan, B. Zhang, D. Tan and J. Zhang, *J. Mater. Chem. A*, 2021, **9**, 14741–14751.
- 138 H. Hu, L. Gui, W. Zhou, J. Sun, J. Xu, Q. Wang, B. He and L. Zhao, *Electrochim. Acta*, 2018, **285**, 70–77.
- 139 B. Kumar, V. Atla, J. P. Brian, S. Kumari, T. Q. Nguyen, M. Sunkara and J. M. Spurgeon, *Angew. Chem., Int. Ed.*, 2017, **56**, 3645–3649.
- 140 Y. Li, B. Wei, M. Zhu, J. Chen, Q. Jiang, B. Yang, Y. Hou, L. Lei, Z. Li, R. Zhang and Y. Lu, *Adv. Mater.*, 2021, **33**, 2102212.
- 141 Q. Hao, H.-x. Zhong, J.-z. Wang, K.-h. Liu, J.-m. Yan, Z.-h. Ren, N. Zhou, X. Zhao, H. Zhang, D.-x. Liu, X. Liu, L.-w. Chen, J. Luo and X.-b. Zhang, *Nat. Synth.*, 2022, **1**, 719–728.
- 142 Y.-N. Gong, C.-Y. Cao, W.-J. Shi, J.-H. Zhang, J.-H. Deng, T.-B. Lu and D.-C. Zhong, *Angew. Chem., Int. Ed.*, 2022, **61**, e202215187.
- 143 X. Zhao, K. Zhao, Y. Liu, Y. Su, S. Chen, H. Yu and X. Quan, *ACS Catal.*, 2022, **12**, 11412–11420.
- 144 M. Feng, X. Wu, H. Cheng, Z. Fan, X. Li, F. Cui, S. Fan, Y. Dai, G. Lei and G. He, *J. Mater. Chem. A*, 2021, **9**, 23817–23827.
- 145 W. Ma, S. Xie, T. Liu, Q. Fan, J. Ye, F. Sun, Z. Jiang, Q. Zhang, J. Cheng and Y. Wang, *Nat. Catal.*, 2020, **3**, 478–487.
- 146 C. Yang, Z. Li, J. Xu, Y. Jiang and W. Zhu, *Green Chem.*, 2024, **26**, 4908–4933.
- 147 Y. Wang, D. Chen, C. Chen and S. Wang, *Acc. Chem. Res.*, 2024, **57**, 247–256.
- 148 C. Chen, X. Zhu, X. Wen, Y. Zhou, L. Zhou, H. Li, L. Tao, Q. Li, S. Du, T. Liu, D. Yan, C. Xie, Y. Zou, Y. Wang, R. Chen, J. Huo, Y. Li, J. Cheng, H. Su, X. Zhao, W. Cheng, Q. Liu, H. Lin, J. Luo, J. Chen, M. Dong, K. Cheng, C. Li and S. Wang, *Nat. Chem.*, 2020, **12**, 717–724.
- 149 W. Huang, T. Bo, S. Zuo, Y. Wang, J. Chen, S. Ould-Chikh, Y. Li, W. Zhou, J. Zhang and H. Zhang, *SusMat*, 2022, **2**, 466–475.
- 150 C. Ling, L. Shi, Y. Ouyang and J. Wang, *Chem. Mater.*, 2016, **28**, 9026–9032.
- 151 G. Gao, A. P. O'Mullane and A. Du, *ACS Catal.*, 2017, **7**, 494–500.
- 152 X. Yang, N. Gao, S. Zhou and J. Zhao, *Phys. Chem. Chem. Phys.*, 2018, **20**, 19390–19397.
- 153 C. Ling, L. Shi, Y. Ouyang, Q. Chen and J. Wang, *Adv. Sci.*, 2016, **3**, 1600180.
- 154 Z. W. Seh, K. D. Fredrickson, B. Anasori, J. Kibsgaard, A. L. Strickler, M. R. Lukatskaya, Y. Gogotsi, T. F. Jaramillo and A. Vojvodic, *ACS Energy Lett.*, 2016, **1**, 589–594.
- 155 M. Alhabeb, K. Maleski, T. S. Mathis, A. Sarycheva, C. B. Hatter, S. Uzun, A. Levitt and Y. Gogotsi, *Angew. Chem., Int. Ed.*, 2018, **57**, 5444–5448.
- 156 M. Naguib, M. Kurtoglu, V. Presser, J. Lu, J. Niu, M. Heon, L. Hultman, Y. Gogotsi and M. W. Barsoum, *Adv. Mater.*, 2011, **23**, 4248–4253.
- 157 A. Thakur, N. Chandran B. S, K. Davidson, A. Bedford, H. Fang, Y. Im, V. Kanduri, B. C. Wyatt, S. K. Nemani, V. Poliukhova, R. Kumar, Z. Fakhraai and B. Anasori, *Small Methods*, 2023, **7**, 2300030.
- 158 P. Urbankowski, B. Anasori, T. Makaryan, D. Er, S. Kota, P. L. Walsh, M. Zhao, V. B. Shenoy, M. W. Barsoum and Y. Gogotsi, *Nanoscale*, 2016, **8**, 11385–11391.
- 159 M. Ghidui, M. Naguib, C. Shi, O. Mashtalir, L. M. Pan, B. Zhang, J. Yang, Y. Gogotsi, S. J. L. Billinge and M. W. Barsoum, *Chem. Commun.*, 2014, **50**, 9517–9520.
- 160 J. Zhu, J. Zhang, R. Lin, B. Fu, C. Song, W. Shang, P. Tao and T. Deng, *Chem. Commun.*, 2021, **57**, 12611–12614.
- 161 A. E. Ghazaly, H. Ahmed, A. R. Rezk, J. Halim, P. O. Å. Persson, L. Y. Yeo and J. Rosen, *ACS Nano*, 2021, **15**, 4287–4293.
- 162 Z. Sun, M. Yuan, L. Lin, H. Yang, C. Nan, H. Li, G. Sun and X. Yang, *ACS Mater. Lett.*, 2019, **1**, 628–632.
- 163 I. T. McCrum and M. T. M. Koper, *Nat. Energy*, 2020, **5**, 891–899.
- 164 P. G. Ghanekar, S. Deshpande and J. Greeley, *Nat. Commun.*, 2022, **13**, 5788.
- 165 K. Bang, D. Hong, Y. Park, D. Kim, S. S. Han and H. M. Lee, *Nat. Commun.*, 2023, **14**, 3004.
- 166 S. Han, S. Lysgaard, T. Vegge and H. A. Hansen, *npj Comput. Mater.*, 2023, **9**, 139.
- 167 Z. Li, F. Chen, W. Bian, B. Kou, Q. Wang, L. Guo, T. Jin, Q. Tang and B. Pan, *Electrochim. Acta*, 2021, **386**, 138465.

