


 Cite this: *Phys. Chem. Chem. Phys.*, 2023, 25, 20405

Protomers of the green and cyan fluorescent protein chromophores investigated using action spectroscopy†

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The photophysics of biochromophore ions often depends on the isomeric or protomeric distribution, yet this distribution, and the individual isomer contributions to an action spectrum, can be difficult to quantify. Here, we use two separate photodissociation action spectroscopy instruments to record electronic spectra for protonated forms of the green (pHBDI⁺) and cyan (Cyan⁺) fluorescent protein chromophores. One instrument allows for cryogenic ($T = 40 \pm 10$ K) cooling of the ions, while the other offers the ability to perform protomer-selective photodissociation spectroscopy. We show that both chromophores are generated as two protomers when using electrospray ionisation, and that the protomers have partially overlapping absorption profiles associated with the $S_1 \leftarrow S_0$ transition. The action spectra for both species span the 340–460 nm range, although the spectral onset for the pHBDI⁺ protomer with the proton residing on the carbonyl oxygen is red-shifted by ≈ 40 nm relative to the lower-energy imine protomer. Similarly, the imine and carbonyl protomers are the lowest energy forms of Cyan⁺, with the main band for the carbonyl protomer red-shifted by ≈ 60 nm relative to the lower-energy imine protomer. The present strategy for investigating protomers can be applied to a wide range of other biochromophore ions.

 Received 8th June 2023,
 Accepted 13th July 2023

DOI: 10.1039/d3cp02661b

rsc.li/pccp

1 Introduction

Green fluorescent protein (GFP), which was first isolated from the *Aequorea victoria* jellyfish over 50 years ago,¹ has become a cornerstone in photobiology.² GFP has revolutionised fluorescence imaging owing to the ease with which it can be incorporated into biological systems and because of its desirable optical properties.³ The photophysics of GFP are governed by a deprotonated *p*-hydroxybenzylidene-2,3-dimethylimidazolinone, pHBDI unit – Fig. 1(a), situated within the protein's β -barrel.⁴ Following the discovery of GFP, a range of derivative fluorescent proteins have been characterised and synthesised, offering tuned absorption and emission profiles, on-off switching, and optical highlighting.^{5–8}

One such variant, the cyan fluorescent protein (CFP), is formed through the substitution of a tyrosine residue in GFP with a tryptophan unit.⁹ This mutation blue-shifts the fluorescence profile due to the combination of an altered chromophore structure 4-((1*H*-indol-3-yl)methylene)-1,2-dimethyl-1*H*-imidazol-5(4*H*)-one (Fig. 1(b)), and a chromophore-specific binding pocket in the protein. Although these chromophores exist in a deprotonated state in the protein, spectroscopic data on the cationic (protonated) forms are useful for calibrating theory and providing

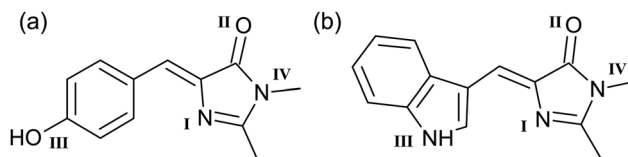


Fig. 1 Structure of (a) pHBDI and (b) Cyan. The protonated cations are denoted pHBDI⁺ and Cyan⁺, with the four possible protonation sites identified. In both cases, protonation site I (imine) corresponds to the lowest energy protomer. Each protomer may have *E* and *Z* geometric configurations about the double bond on the methine bridge. In all cases, the *Z* isomer is lower in energy. The *E* isomer, if present, is expected in only trace quantities.

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 † Electronic supplementary information (ESI) available: ¹H NMR spectrum for Cyan; Cartesian coordinates of optimised geometries. See DOI: <https://doi.org/10.1039/d3cp02661b>

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benchmark data for understanding spectra for protonation equilibria in solution.^{10,11}

In solution and in the gas phase, the protonated forms of these chromophores, *p*HBDI⁺ and Cyan⁺, can exist as several protomers (protonation isomers), which may exhibit distinct photophysical properties.¹² Characterisation of protomers (or deprotomers for anions) and probing the impact of protonation site on molecular photophysics is a topical area in gas-phase spectroscopy due to the emerging analytical capabilities to separate and probe protomers;¹³ such selectivity is almost always impossible in solution due to rapid proton exchange, with solution experiments measuring some solvent-dependent average.

There are several strategies for inferring the presence of coexisting protomers in gas-phase experiments, including by comparing infrared and ultraviolet action spectroscopy measurements with electronic structure calculations, such as done for benzocaine,¹⁴ DNA bases,¹⁵ and flavins.¹⁶ Similar spectroscopic identifications of coexisting protomers have been made at cryogenic temperatures through analysis of vibronic structure.^{17,18} Often, however, protomers have overlapping absorption profiles but may have distinct excited-state dynamics, making them difficult to study confidently without isomer-selective action spectroscopy techniques. As an example, protomers of 4-dicyanomethylene-2-methyl-6-*para*-dimethylaminostyryl-4*H*-pyran (DCM) dye molecules have been separated using ion mobility spectrometry and were shown to have well-resolved action spectra (shifted by nearly 200 nm) and protomer-specific photoisomerisation dynamics.¹² Isomer selectivity in action spectroscopy experiments is most commonly achieved by incorporating ion mobility as a separation dimension.^{13,19} While the occurrence of coexisting protomers is becoming a common theme in gas-phase spectroscopy, general trends or ‘rules of thumb’ for the preponderance of protomers can be difficult to determine because, aside from the electrospray solvent, relative protomer yields can depend on a vast number of factors including solution pH, electrospray geometry and needle voltage, desolvation conditions, and collisional treatment of ions as they are introduced into vacuum.^{20,21} Additionally, because most gas-phase action spectroscopy instruments are custom designed, it can be difficult to know if the protomer distribution generated in one experiment is the same as for another.

Here, we report photodissociation action spectra for *p*HBDI⁺ and Cyan⁺, produced using electrospray ionisation. While the gas-phase deprotonated anions of *p*HBDI chromophores have been studied extensively (see ref. 22 for a recent summary of relevant literature) and deprotonated Cyan has been studied with photoelectron spectroscopy,²³ the protonated cations have been studied at room-temperature only,^{24–26} with limited discussion of protomers. In this work, photodissociation action spectra of *p*HBDI⁺ and Cyan⁺ were recorded using two instruments, the first offering the capabilities for measurements at $T \approx 300$ K and $T = 40 \pm 10$ K, and the second offering isomer (protomer) selectivity, although the ions are heated to $T = 400$ – 500 K during the isomer selection step. The results of both sets of experiments are consistent with two coexisting protomers, possessing overlapping action spectra in the visible. The protomers were

assigned based on electronic structure calculations of their stabilities and Franck–Condon simulation of absorption profiles.

2 Experimental

The *p*HBDI molecule used in this work was available from an earlier study.²² The Cyan molecule was synthesised in a two-step reaction according to an established procedure involving the preliminary synthesis of an azalactone followed by its aminolysis with methylamine.²⁷ The structures and purity of the compounds were confirmed by ¹H spectroscopy – see ESI.†

2.1 Room-temperature and cryogenic photodissociation action spectroscopy

Photodissociation action spectra at $T \approx 300$ K and $T = 40 \pm 10$ K were recorded, at the PIIM laboratory in Marseille, France, by trapping electrosprayed ions in a cryogenically-cooled quadrupole ion trap (QIT) and irradiating the ions with tunable laser light. Complete details of the apparatus are given elsewhere.^{28,29} Briefly, *p*HBDI⁺ and Cyan⁺ were generated through electrospray ionisation of a millimolar solution of the target molecule dissolved in a 1 : 1 (v/v%) methanol–water mixture. Several drops of concentrated acetic acid were added to the solutions to protonate the target molecule. After electrospray, ions were stored in an octopole for ≈ 100 ms before being extracted and transferred into the QIT. Helium buffer gas was injected into the QIT using a Parker pulsed valve, serving to collisionally decelerate and thermalise the ions to a temperature of $T \approx 300$ K or $T = 40 \pm 10$ K.³⁰ The cryogenic temperature was achieved by cooling of the trap with a Leybold helium-cooled cryohead to $T \approx 20$ K (several temperature sensors across the ion trap assembly vary between $T = 12$ K to 35 K).³⁰ An auxiliary radio frequency was applied to the trap and tuned to the resonance frequencies of the fragment ions in order to minimise collision-induced fragment signal.³¹ Stored ions were irradiated in the trap using light from an optical parametric oscillator laser (OPO, EKSPILA NT342B), operating at 10 Hz repetition rate with ≈ 10 ns pulse width and ≈ 10 cm⁻¹ bandwidth. The laser fluence (2 – 4 mJ pulse⁻¹ for $\lambda > 400$ nm and 4 – 7 mJ pulse⁻¹ for $\lambda < 400$ nm) was constantly measured during the spectral acquisitions. After irradiation, photofragment and parent ions were extracted from the QIT by applying a negative pulse (-400 V) on the exit cap electrode. The ions were then accelerated, by a second negative pulse (-2800 V) applied to a Gauss tube, into the field free region of a linear time-of-flight mass spectrometer, where they were detected by microchannel plate assembly. For *p*HBDI⁺, the width of the pulse applied to the Gauss tube was optimised to accelerate both parent ions and small photofragments. For Cyan⁺, this condition could not be achieved due to the limited m/z range accessible in the experiment; consequently, photodepletion spectroscopy of the parent Cyan⁺ was performed. Action spectra were derived by plotting laser-on photofragment or photodepletion yield normalised with laser-off signal and laser fluence.



2.2 Protomer-specific photodissociation action spectroscopy

Protomer-specific photodissociation action spectra for electro-sprayed $p\text{HBDI}^+$ and Cyan^+ were recorded using a modified instrument combining differential mobility spectrometry (DMS) with mass spectrometry (described in detail in ref. 32–35), at the University of Waterloo, Canada. The target molecules were dissolved in a 50:50 water:methanol solution ($\approx 600 \text{ ng mL}^{-1}$) with 0.1% formic acid to assist protonation. Following electrospray, the protonated cations were carried, by dry N_2 gas heated to $\approx 150 \text{ }^\circ\text{C}$ doped with 1.5 mol% acetonitrile ($p\text{HBDI}^+$), 1.5 mol% propan-2-ol (Cyan^+) or no dopant ($p\text{HBDI}^-$), through the planar DMS cell. The addition of a dopant (mobility modifier) to the carrier gas causes the formation of microsolvated clusters, which, dependent on the protic nature of the dopant, dictates the yield and interconversion of protomers in the DMS cell.^{36–38} Isomer population ionograms were acquired by applying a separation voltage (SV) across the DMS cell (SV = 3500 V for $p\text{HBDI}^+$ and 2900 V for Cyan^+) and scanning the compensation voltage (CV). For the photodissociation measurements, cations transmitted at a fixed CV were passed into a triple-quadrupole mass spectrometer (Q1–Q3), where Q3 can be operated as a traditional quadrupole filter or a linear ion trap. These ions were mass selected in Q1 (m/z 217 for $p\text{HBDI}^+$ and m/z 240 for Cyan^+), transmitted through Q2, and accumulated in Q3. The accumulated ion bunches were irradiated with light from an OPO (Horizon II, Continuum, loosely focused, 4–7 mJ pulse⁻¹) pumped by a pulsed Nd:YAG laser (10 Hz, Surelite, Continuum). Parent and photofragment ions were ejected from Q3 and counted with a channeltron ion detector. Action spectra were derived by plotting laser-on photofragment or photodepletion yield normalised with respect to laser-off signal and laser-fluence. The error in photofragmentation efficiency, for each wavelength, was obtained by propagation of the error in photofragmentation efficiency, assessed over 25 laser interrogation/ejection cycles within Q3, and the variation in laser power assessed over 100 shots.

The predominant photofragment from both protomers of $p\text{HBDI}^+$ was m/z 56 ($\text{C}_3\text{H}_6\text{N}$), with the next four most abundant ions at m/z 160, 186, 119, and 91. For Cyan^+ , the predominant photofragment from both protomers was m/z 56 ($\text{C}_3\text{H}_6\text{N}$) with the next four most abundant ions at m/z 157, 183, 130, and 115. There were no obvious marker m/z products from a given protomer ensemble that clearly reflect the protonation site, presumably because dissociation is statistical in nature and rearrangement barriers are lower in energy than dissociation barriers.³⁹

2.3 Computational

Electronic structure calculations were performed using the Gaussian 16B.01⁴⁰ and ORCA 5.0.3 software packages.^{41,42} Geometry optimisations and vibrational frequencies for the S_0 and S_1 electronic states of neutral and cationic forms of the chromophores were determined at the $\omega\text{B97X-D/aug-cc-pVDZ}$ level of theory.^{43–45} Single-point energies were computed at the CCSD(T)/aug-cc-pVDZ and DLPNO-CCSD(T)/aug-cc-pVDZ levels of theory. Vertical excitation energies were computed at the

DLPNO-STEOM-CCSD/aug-cc-pVDZ level of theory.^{46,47} The DLPNO-CCSD(T) method has been benchmarked to calculate relative energies to within $\approx 20 \text{ meV}$ ($\approx 2 \text{ kJ mol}^{-1}$) of established ref. 48 and 49, while the DLPNO-STEOM-CCSD method has previously reproduced vertical transition energies for a series of $p\text{HBDI}$ -based anions and other biochromophores to within $\approx 0.1 \text{ eV}$.^{22,50}

Franck–Condon (FC) simulations of the absorption spectra (at $T = 20 \text{ K}$ – there was no substantial change at $T = 40 \text{ K}$), with the exception of Cyan^+ protomer **I**, were performed using ezSpectrum 3.0 (ezFCF).⁵¹ The FC factors were obtained using the parallel mode approximation with a maximum of four quanta of vibrations in the target state. The final spectra were constructed by convolving the simulated vibronic transitions with Gaussian functions (10 cm^{-1} FWHM) approximating the laser bandwidth. Similar Franck–Condon simulations can satisfactorily describe photoelectron spectra of cryogenic $p\text{HBDI}^-$.⁵² The simulation for Cyan^+ protomer **I**, also convolved with 10 cm^{-1} FWHM Gaussian functions, used the vertical Hessian approximation in Gaussian 16,⁵³ due to a large change in geometry associated with torsion of the methine bridge between the ground and excited states.

3 Results and discussion

3.1 $p\text{HBDI}^+$

Photodissociation action spectra recorded for $p\text{HBDI}^+$ at $T \approx 300 \text{ K}$ and $T = 40 \pm 10 \text{ K}$ are shown in Fig. 2(a) and (b), respectively. At $T \approx 300 \text{ K}$, the spectrum spans the 340–460 nm range with the peak spectral response at $410 \pm 1 \text{ nm}$. The wavelength of maximum response in the cryogenic spectrum is $407 \pm 1 \text{ nm}$, which is blue-shifted by $\approx 3 \text{ nm}$ compared with the $T \approx 300 \text{ K}$ spectrum. At cryogenic temperature, the spectrum sharpens and an additional peak centred at $428 \pm 1 \text{ nm}$ is resolved. We immediately note that these spectra are inconsistent with the two earlier reports (at $T \approx 300 \text{ K}$), which were recorded by monitoring photo-induced neutral fragment production²⁴ and photofragments generated in an ion trap²⁶ – both spectra are included in Fig. 2(a) for comparison. While these two earlier measurements are in good agreement with each other and the wavelength of maximum response is consistent with the current work, the earlier spectra are substantially narrower and do not show the red-edge shoulder evident in the current work.

The ionogram of electro-sprayed $p\text{HBDI}^+$ recorded using DMS is shown in Fig. 2(e). Two isomeric (protomer) populations are discernible, centered at CVs of -26.3 V (**I**, blue) and -21.5 V (**II**, red), with a 3 : 1 abundance. The asymmetry in peak **II** is probably a signal-to-noise artefact. Protomer-specific photodissociation action spectra recorded for mobility-selected ions at the two CVs are given in Fig. 2(f) and (g), respectively. Both spectra have similar band maxima ($\approx 405 \text{ nm}$) and overall shape, although the spectrum for protomer **I** is narrower than for protomer **II**, in part due to the presence of a shoulder at $\approx 440 \text{ nm}$ for protomer **II**. The protomer-specific photodissociation maxima are in close



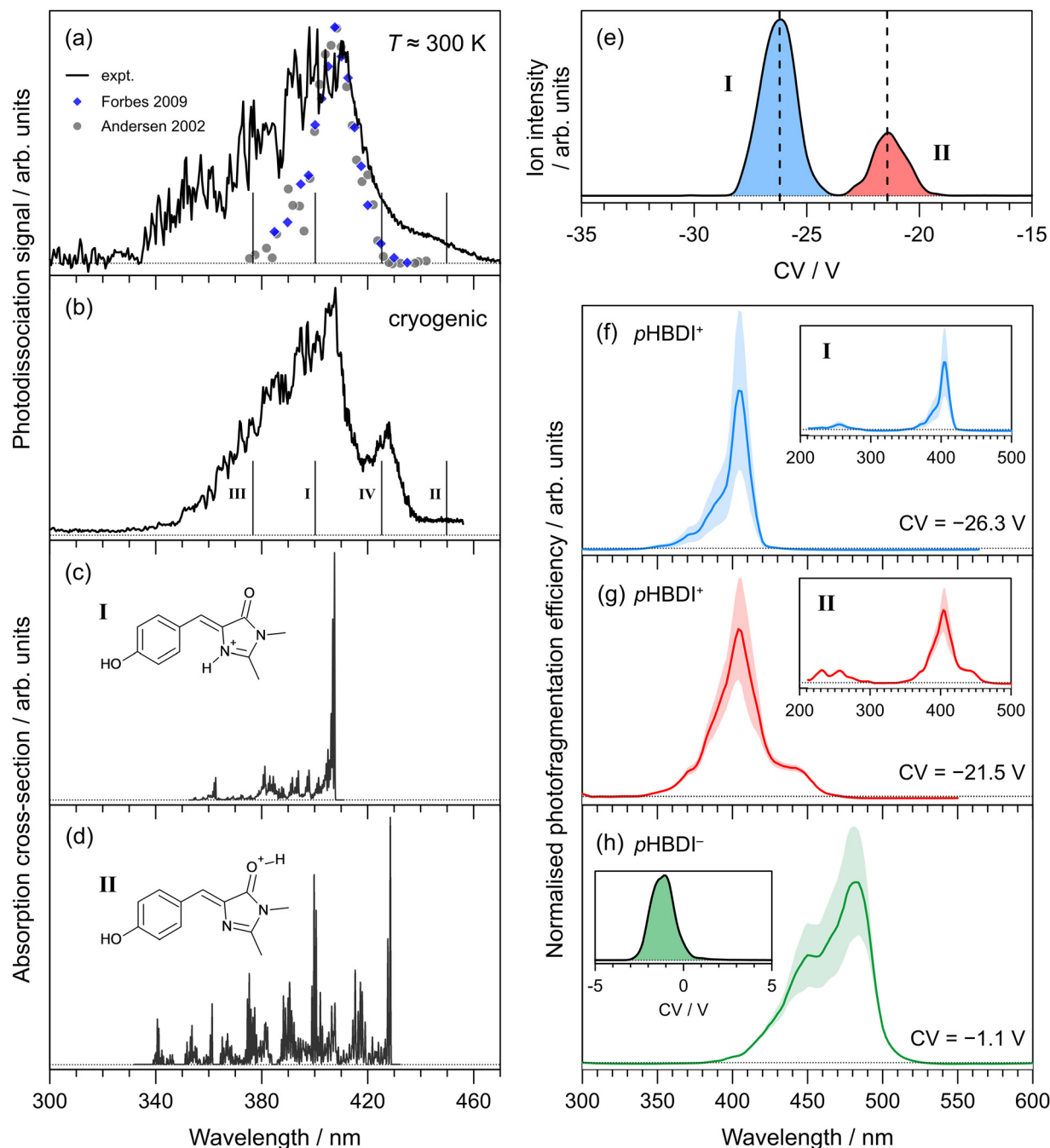


Fig. 2 Action spectroscopy of *p*HBDI⁺: (a) photodissociation spectrum at $T \approx 300$ K [grey dots taken from ref. 24 and blue diamonds from ref. 26], (b) photodissociation spectrum at cryogenic temperature. Vertical bars indicate the calculated vertical excitation wavelengths for protomers I–IV, determined at the DLPNO-STEOM-CCSD(T)/aug-cc-pVDZ level of theory. Franck–Condon simulation of the $S_1 \leftarrow S_0$ absorption profiles (at $T = 20$ K) for: (c) protomer I (lowest energy protomer) assigned as the imine protomer I in Fig. 1, and (d) protomer II (second lowest energy protomer) assigned as the carbonyl protomer II in Fig. 1. Protomer-specific photodissociation action spectroscopy: (e) ionogram recorded using DMS heated to $T = 150$ °C with 1.5 mol% acetonitrile entrained in N_2 at SV = 3500 V showing two isomer populations, assigned to protomers I and II, (f) photodissociation spectrum for protomer I, (g) photodissociation spectrum for protomer II. Panel (h) shows the photodissociation spectrum recorded for (deprotonated) *p*HBDI⁻ and ionogram (inset), consistent with a single deprotomer [SV = 3000 V, pure N_2 carrier gas], and with earlier spectra.^{22,54} Dotted horizontal baselines indicate the zero signal level.

agreement with that for the $T \approx 300$ K spectrum in Fig. 2(a), and it is evident that some combination of the protomer-specific photodissociation spectra can describe the $T \approx 300$ K and cryogenic spectra. Furthermore, the action spectrum for

protomer I is in good agreement with the earlier ion storage ring and ion trap photodissociation reports (Fig. 2(a)), supporting that the conditions in the earlier experiments produced (predominantly) a single protomer.



To help assign the action spectra to individual protomers, we performed electronic structure calculations on the possible protonated structures identified in Fig. 1(a). Relative energies of each protomer are summarised in Table 1, showing that protomer **I**, with protonation on the imine nitrogen, is the most stable form. The second lowest energy form, protomer **II**, lies 0.38 eV higher in energy and corresponds to protonation on the carbonyl oxygen. The other protomers, **III** and **IV**, lie much higher in energy and are unlikely to be formed during electrospray ionisation. Calculated vertical excitation wavelengths (DLPNO-STEOM-CCSD/aug-cc-pVDZ level of theory) for the protomers are indicated as vertical bars in Fig. 2(a) and (b) (values given in Table 1). Earlier benchmarking calculations on *p*HBDI-based deprotonated anions with the same level of theory determined that vertical excitation energies, on average, are within ≈ 0.05 eV (≈ 15 nm in the present spectral range) of experiment.²² The calculated vertical excitation wavelength for protomer **I** is within 10 nm of the action spectrum peak. On the other hand, the calculated value for protomer **II** is red-shifted by ≈ 20 nm compared with the longest-wavelength shoulder feature in the cryogenic action spectrum.

Franck-Condon simulation of the $S_1 \leftarrow S_0$ absorption spectra for protomers **I** and **II**, performed at the ω B97X-D/aug-cc-pVDZ level of theory, are shown in Fig. 2(c) and (d), respectively, and have been translated to coincide with the experimental spectra. The simulation for protomer **I** is consistent with the mobility-selected spectrum for protomer **I** as well as the earlier $T \approx 300$ K measurements (Fig. 2(a)), and describes the most intense portion of our cryogenic spectrum (Fig. 2(b)). The Franck-Condon simulation for protomer **II** is consistent with the additional red-edge shoulder feature observed in our $T \approx 300$ K and cryogenic action spectra, and the mobility-selected action spectrum for protomer **II**. The extended progression character of the simulated band also describes the broadened tail on the blue-edge of those experimental spectra.

We note that *p*HBDI⁺ (and Cyan⁺) can exist as *E* and *Z* isomers with respect to the double bond on the methine bridge. Isomer-selective action spectroscopy experiments using a drift-

tube instrument on the deprotonated anion, *p*HBDI⁻,⁵⁴ have shown that the relative abundance of the *E* and *Z* isomers depends on collisional treatment of the ions as they are introduced into vacuum. However, the mobility-selected action spectra for the *E* and *Z* isomers of *p*HBDI⁻ are essentially identical (*i.e.* indistinguishable), consistent with vertical excitation wavelength calculations. In the present DMS experiments, only a single ionogram peak was observed for *p*HBDI⁻. For *p*HBDI⁺ and Cyan⁺, we expect the action spectra to predominately reflect the *Z* isomers for a given protomer because the energies of the *E*-isomers are substantially higher than for the *Z*-isomers (Table 1). Furthermore, electronic structure calculations of vertical excitation wavelengths for the *E* isomers of the imine protomers are within a few nanometers of those for the *Z* isomers.

It is interesting to note that the mobility-selected action spectra for protomers **I** and **II** have the same wavelengths of maximum response. There are two possible explanations. First, this is coincidence and indeed the mobility-selective experiments probe pure ensembles of each protomer due to clear protomer separation (Fig. 2(e)). Second, there may be some degree of conversion of protomer **II** into the more stable protomer **I**, which occurs after DMS isomer separation and while the ions are being transferred into the trap. Such an interconversion may occur due to the temperature of the ions ($T \approx 500$ – 600 K) and collisions associated with focussing and RF electric field as ions are transferred into the higher vacuum environment. Whatever the case, it is clear that the action spectra for *p*HBDI⁺ recorded in two separate instruments, incorporating different electrospray ionisation sources and ion optics, show the co-existence of the two lowest energy protomers. Conversely, the earlier ion storage ring and ion trap photodissociation spectra appear consistent with a single protomer.

We conclude that *p*HBDI⁺ is generated as imine and carbonyl protomers in both the Marseille and Waterloo photodissociation experiments.

3.2 Cyan⁺

The photodissociation action spectrum recorded for Cyan⁺ at $T \approx 300$ K is shown in Fig. 3(a). At $T \approx 300$ K, the spectrum spans the 340–460 nm range, with the peak spectral response at 412 ± 2 nm. Our spectrum is substantially different to an earlier report (filled circles in Fig. 3(a)), which was recorded in an ion storage ring by monitoring the production of neutral photofragments.²⁵ The literature spectrum peak is red-shifted by ≈ 40 nm and spans a slightly narrower spectral range (420–510 nm). A rough cryogenic photodepletion spectrum for Cyan⁺ recorded in the Marseille experiment is included in Fig. 3(a), and is substantially narrower than the $T \approx 300$ K spectrum and with a slightly (≈ 10 nm) blue-shifted maximum.

In a similar vein to *p*HBDI⁺, we computed the lowest energy protomers (Table 1), finding that protonation on the imine nitrogen (protomer **I**) and carbonyl group (protomer **II**) corresponds to the two lowest energy protomers, respectively. The other protomers, **III** and **IV**, lie substantially higher in energy and are unlikely to be formed during electrospray ionisation. Calculated vertical excitation wavelengths (DLPNO-STEOM-

Table 1 Calculated protomer energetics for *p*HBDI⁺ and Cyan⁺ (see Fig. 1 for the protomer numbering scheme). ΔE is the difference in total energy, relative to protomer **I** for each species, computed at the DLPNO-CCSD(T)/aug-cc-pVDZ level of theory. In parentheses are the differences in energy (in kJ mol⁻¹) between the *Z* and *E* isomers of protomers **I** and **II**. The vertical excitation energy (VEE) was computed at the DLPNO-STEOM-CCSD/aug-cc-pVDZ level of theory

Protomer	<i>p</i> HBDI ⁺		Cyan ⁺	
	ΔE^a /eV	VEE/eV	ΔE^a /eV	VEE/eV
I	0.00 (14) ^b	3.10	0.00 (31) ^b	3.00
II	0.38 (24) ^b	2.76	0.38 (41) ^b	2.75
III	2.04	3.29	1.26	3.37
IV	1.20	2.92	1.23	2.93

^a For protomers **I**, **II**, and **III**, calculations at the CCSD(T)/aug-cc-pVDZ level of theory gave the same ΔE values, supporting the use of the less demanding DLPNO-CCSD(T) method. ^b These values are greater than the 11 kJ mol⁻¹ energy difference between *Z*- and *E*-*p*HBDI⁻.⁵⁴ It is unlikely that these *E* isomers will form considering the substantial energy differences.



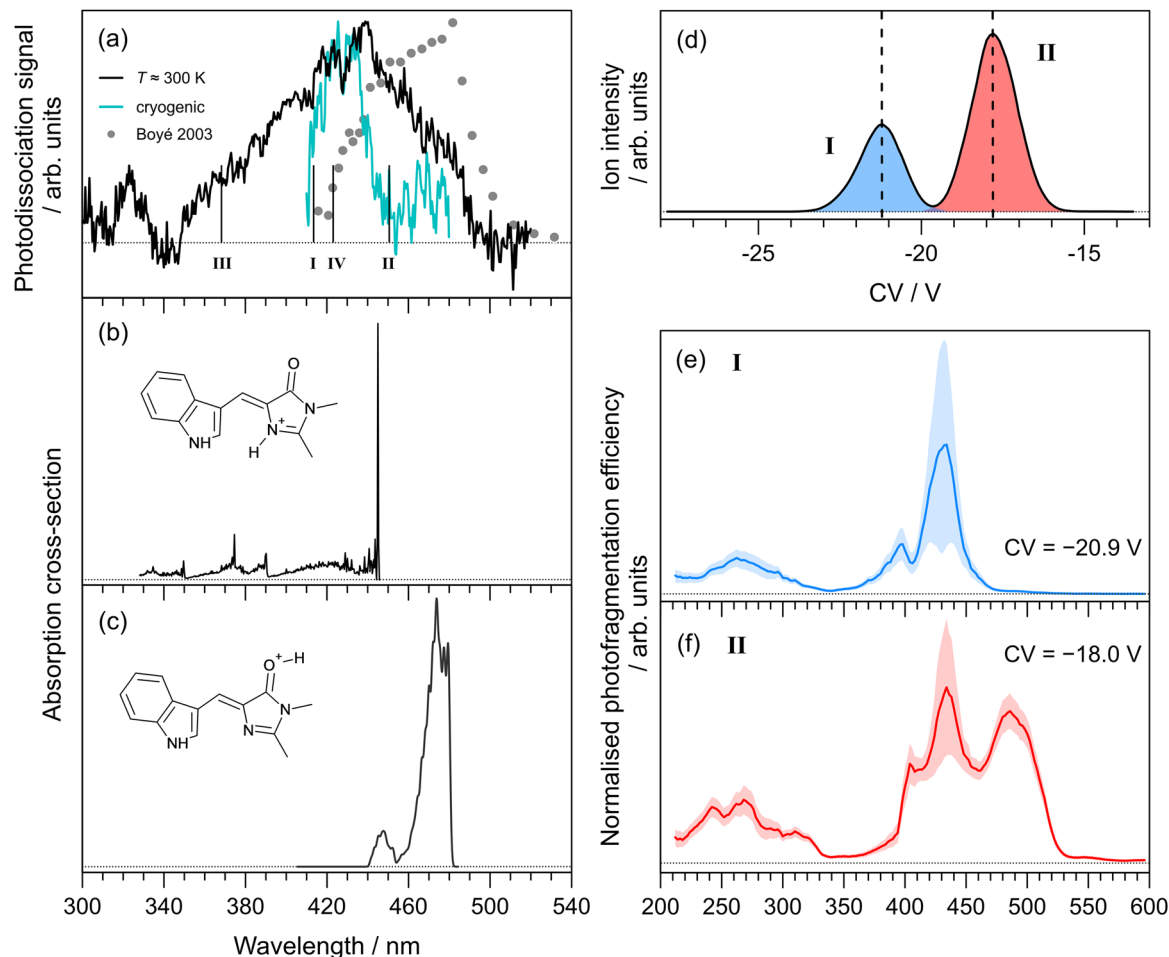


Fig. 3 Action spectroscopy of Cyan⁺: (a) photodissociation spectrum at $T \approx 300$ K [dots are from ref. 25]. Vertical bars indicate the calculated vertical excitation wavelengths for protomers I–IV, determined at the DLPNO-STEOM-CCSD(T)/aug-cc-pVDZ level of theory. Franck–Condon simulation of the $S_1 \leftarrow S_0$ absorption profiles (at $T \approx 20$ K) for: (b) the lowest energy protomer, I, and (c) the second lowest energy protomer, II. Protomer-specific photodissociation spectroscopy: (d) ionogram recorded using DMS heated to $T = 150$ °C with 1.5 mol% propan-2-ol entrained in N₂ at SV = 2900 V showing two isomer populations assigned to protomers I and II, (e) photodissociation spectrum for species I, (f) photodissociation spectrum for species II.

CCSD/aug-cc-pVDZ level of theory) for the possible protomers are indicated as vertical bars in Fig. 3(a) (values given in Table 1). Taking into account the precision of the computational method (± 0.2 eV), the two lowest energy protomers could satisfactorily describe the spectra.

Franck–Condon simulation of the $S_1 \leftarrow S_0$ absorption profiles (translated to coincide with the experimental spectra) for the imine and carbonyl protomers are shown in Fig. 3(b) and (c), respectively. The imine protomer was difficult to model because of a large internal rotation of single bond as part of the methine bridge (S_0 at 21.5 and S_1 at 87.7), although the vertical excitation is bright (oscillator strength, $f = 0.60$). We assign the main part our $T \approx 300$ K spectrum to protomer I, and the red-edge (as well as the literature spectrum) to protomer II. These assignments are also based on the protomer-specific spectra described below.

The ionogram for electrosprayed Cyan⁺ recorded using the Waterloo instrument is shown in Fig. 3(d). As with *p*HBDI⁺, two isomeric (protomer) populations are discernible, centred at CVs of -20.9 V (I, blue) and -18.0 V (II, red), with approximately a

2:1 abundance. Protomer-specific photodissociation action spectra recorded for mobility-selected ions at the two CVs are given in Fig. 3(d) and (e), respectively. The photodissociation spectrum for protomer I has maximum response at 434 ± 3 nm, while protomer II has the additional peak at 488 ± 3 nm. These assignments are consistent with the energy ordering of the calculated the vertical excitation wavelengths.

It is significant to note that the longer wavelength peak at ≈ 430 nm in the protomer-specific spectrum assigned to protomer II is not reproduced in the Franck–Condon simulation for the $S_1 \leftarrow S_0$ transition in Fig. 3(c). Furthermore, our DLPNO-STEOM-CCSD/aug-cc-pVDZ calculations indicated no higher-lying excited state with a substantial oscillator strength that could account for this feature. We therefore hypothesise that this spectral feature is associated with conversion of protomer II to protomer I after DMS separation, presumably due to solvent clustering combined with a low barrier to proton shuffling.^{38,55} This conversion is consistent with the fact that protomer II is higher in energy and, on statistical thermodynamic grounds, would be expected to convert to protomer I



when energised. The same arguments can be applied to $p\text{HBDI}^+$ protomer **II**. In principle, such protomer interconversions might be minimised by trialling DMS separation with a range of non-protic dopants.

In summary, as for $p\text{HBDI}^+$, we conclude that Cyan^+ is generated as imine and carbonyl protomers in both the Marseille and Waterloo photodissociation experiments.

4 Conclusions

This work has demonstrated that electrospray ionisation of protonated $p\text{HBDI}$ and Cyan chromophores in two disparate electrospray ionisation experiments each generates two protomers, and has demonstrated the capacity for DMS coupled with laser spectroscopy to quantify the protomers and record protomer-specific action spectra. For both chromophores, protonation on the imine nitrogen is the lowest energy form, followed by protonation on the carbonyl group. Additional protomers could, in principle, be generated in other experiments due to the large number of factors that contribute to protomer abundance. Because the protomer-specific photodissociation spectra for a given chromophore overlap, it is difficult to disentangle relative protomer abundances or contributions in experiments that do not offer protomer separation. Future work will seek to use ion mobility spectrometry coupled with other laser spectroscopies, including photoisomerisation action spectroscopy⁵⁶ and femtosecond time-resolved strategies,⁵⁷ to probe the photochemistry of each protomer individually.

The occurrence of co-existing gas-phase protomers (or deprotomers in negative-mode electrospray ionisation) appears a common issue for many biochromophores,^{58–60} and will likely become an increasingly important as action spectroscopy experiments are applied to more complex biomolecules with many heteroatom sites (e.g. N and O atoms), allowing for a large number of protomers. Robust action spectroscopy studies on such species should have provisions for protomer selectivity.

The spectra presented in this paper complement recent DMS deprotomer-selected photodissociation action spectra on the Kaede chromophore,⁵⁵ and, similarly, required a dopant to assist DMS separation. As for $p\text{HBDI}^+$ and Cyan^+ , it was found that the higher energy deprotomer probably contained a signature of the lower energy deprotomer. Further quantitative information on the propensity and barriers for protomer/deprotomer interconversion could be obtained through collision-induced tandem ion mobility spectrometry experiments,^{12,58,59,61} where precursor species are selected, collisionally activated, and the products then separated and quantified. Such experiments could be conducted as a function of the dopant molecule to investigate how properties, such as hydrogen-bonding propensity of the dopant,^{36–38} affect isomerisation barriers.

Data availability

The data that support the findings of this study are available from the corresponding author upon reasonable request.

Author contributions

Photodissociation experiments at $T \approx 300$ K and $T = 40 \pm 10$ K (cryogenic) were performed by JD, JAN, EKA, and JNB in the laboratory of JAN. Protomer-specific action spectroscopy experiments were performed by CRMR and CI in the laboratory of WSH. Quantum chemical calculations were performed by EKA, and Franck–Condon simulations were performed by JD and JNB. The Cyan molecule was synthesised by IC and ANC. The manuscript was prepared by EKA and JNB, and discussed by all authors. The project was managed by JNB and MHS.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

JNB acknowledges funding from an EPSRC New Investigator Award (EP/W018691). JAN acknowledges funding from the Agence Nationale de la Recherche (ANR, HYDRAE project ANR-21-CE30-0004-01) and additional support from the French Programme National “Physique et Chimie du Milieu Interstellaire” (PCMI) of the CNRS/INSU with the INC/INP, co-funded by the CEA and the CNES. WSH acknowledges financial support provided by the Natural Sciences and Engineering Research Council (NSERC) of Canada in the form of Discovery and Alliance grants, and Mitacs in the form of an Accelerate grant. EKA thanks the University of East Anglia for a doctoral studentship. Travel funding was provided by the Swedish Foundation for International Cooperation in Research and Higher Education (STINT) Grant for Internationalisation programme (PT2017-7328 to MHS and JNB). Electronic structure calculations were carried out on the High Performance Computing Cluster supported by the Research and Specialist Computing Support service at the University of East Anglia.

Notes and references

- 1 R. Y. Tsien, *Annu. Rev. Biochem.*, 1998, **67**, 509–544.
- 2 *Fundamentals of Fluorescence Imaging*, ed. G. Cox, Jenny Stanford Publishing, 2019.
- 3 N. C. Shaner, P. A. Steinbach and R. Y. Tsien, *Nat. Methods*, 2005, **2**, 905–909.
- 4 M. Ormö, A. B. Cubitt, K. Kallio, L. A. Gross, R. Y. Tsien and S. J. Remington, *Science*, 1996, **273**, 1392–1395.
- 5 G.-J. Kremers, J. Goedhart, E. B. van Munster and T. W. J. Gadella, *Biochemistry*, 2006, **45**, 6570–6580.
- 6 G. D. Malo, L. J. Pouwels, M. Wang, A. Weichsel, W. R. Montfort, M. A. Rizzo, D. W. Piston and R. M. Wachter, *Biochemistry*, 2007, **46**, 9865–9873.
- 7 B. Seefeldt, R. Kasper, T. Seidel, P. Tinnefeld, K.-J. Dietz, M. Heilemann and M. Sauer, *J. Biophotonics*, 2008, **1**, 74–82.
- 8 J. S. Paige, K. Y. Wu and S. R. Jaffrey, *Science*, 2011, **333**, 642–646.



- 9 R. Heim, D. C. Prasher and R. Y. Tsien, *Proc. Natl. Acad. Sci. U. S. A.*, 1994, **91**, 12501–12504.
- 10 X. He, A. F. Bell and P. J. Tonge, *FEBS Lett.*, 2003, **549**, 35–38.
- 11 S. P. Laptinok, J. Conyard, P. C. B. Page, Y. Chan, M. You, S. R. Jaffrey and S. R. Meech, *Chem. Sci.*, 2016, **7**, 5747–5752.
- 12 J. N. Bull, N. J. A. Coughlan and E. J. Bieske, *J. Phys. Chem. A*, 2017, **121**, 6021–6027.
- 13 S. J. P. Marlton and A. J. Trevitt, *Chem. Commun.*, 2022, **58**, 9451–9467.
- 14 S. Warnke, J. Seo, J. Boschmans, F. Sobott, J. H. Scrivens, C. Bleiholder, M. T. Bowers, S. Gewinner, W. Schöllkopf, K. Pagel and G. von Helden, *J. Am. Chem. Soc.*, 2015, **137**, 4236–4242.
- 15 S. Ø. Pedersen, C. S. Byskov, F. Turecek and S. B. Nielsen, *J. Phys. Chem. A*, 2014, **118**, 4256–4265.
- 16 E. Matthews, R. Cercola and C. Dessent, *Molecules*, 2018, **23**, 2036.
- 17 J. A. Noble, C. Dedonder-Lardeux, J. Mascetti and C. Jouvét, *Chem. – Asian J.*, 2017, **12**, 1523–1531.
- 18 D. Müller and O. Dopfer, *Phys. Chem. Chem. Phys.*, 2020, **22**, 18328–18339.
- 19 H. Zettergren, A. Domaracka, T. Schlathölter, P. Bolognesi, S. Díaz-Tendero, M. Abuda, S. Tosić, S. Maclot, P. Johnsson, A. Steber, D. Tikhonov, M. C. Castrovilli, L. Avaldi, S. Bari, A. R. Milosavljević, A. Palacios, S. Faraji, D. G. Piekarski, P. Rousseau, D. Ascenzi, C. Romanzin, E. Erdmann, M. Alcamí, J. Kopyra, P. Limão-Vieira, J. Kočišek, J. Fedor, S. Albertini, M. Gatchell, H. Cederquist, H. T. Schmidt, E. Gruber, L. H. Andersen, O. Heber, Y. Toker, K. Hansen, J. A. Noble, C. Jouvét, C. Kjær, S. B. Nielsen, E. Carrascosa, J. Bull, A. Candian and A. Petrigiani, *Eur. Phys. J. D*, 2021, **75**, 152.
- 20 A. B. Attygalle, H. Xia and J. Pavlov, *J. Am. Soc. Mass Spectrom.*, 2017, **28**, 1575–1586.
- 21 H. Xia and A. B. Attygalle, *J. Am. Soc. Mass Spectrom.*, 2017, **28**, 2580–2587.
- 22 E. K. Ashworth, M. H. Stockett, C. Kjær, P. C. B. Page, S. R. Meech, S. B. Nielsen and J. N. Bull, *J. Phys. Chem. A*, 2022, **126**, 1158–1167.
- 23 M. A. Parkes, A. Bennett and H. H. Fielding, *Mol. Phys.*, 2019, **117**, 3027–3035.
- 24 L. H. Andersen, A. Lapierre, S. B. Nielsen, I. B. Nielsen, S. U. Pedersen, U. V. Pedersen and S. Tomita, *Eur. Phys. J. D*, 2002, **20**, 597–600.
- 25 S. Boyé, I. B. Nielsen, S. B. Nielsen, H. Krogh, A. Lapierre, H. B. Pedersen, S. U. Pedersen, U. V. Pedersen and L. H. Andersen, *J. Chem. Phys.*, 2003, **119**, 338–345.
- 26 M. W. Forbes and R. A. Jockusch, *J. Am. Chem. Soc.*, 2009, **131**, 17038–17039.
- 27 V. Voliani, R. Bizzarri, R. Nifos, S. Abbruzzetti, E. Grandi, C. Viappiani and F. Beltram, *J. Phys. Chem. B*, 2008, **112**, 10714–10722.
- 28 I. Alata, J. Bert, M. Broquier, C. Dedonder, G. Féraud, G. Grégoire, S. Soorkia, E. Marceca and C. Jouvét, *J. Phys. Chem. A*, 2013, **117**, 4420–4427.
- 29 J. A. Noble, E. Marceca, C. Dedonder, I. Carvin, E. Gloaguen and C. Jouvét, *Eur. Phys. J. D*, 2021, **75**, 1–13.
- 30 N. Esteves-López, C. Dedonder-Lardeux and C. Jouvét, *J. Chem. Phys.*, 2015, **143**, 074303.
- 31 H. Kang, G. Féraud, C. Dedonder-Lardeux and C. Jouvét, *J. Phys. Chem. Lett.*, 2014, **5**, 2760–2764.
- 32 J. L. Campbell, M. Zhu and W. S. Hopkins, *J. Am. Soc. Mass Spectrom.*, 2014, **25**, 1583–1591.
- 33 C. Liu, J. C. Y. L. Blanc, J. Shields, J. S. Janiszewski, C. Ieritano, G. F. Ye, G. F. Hawes, W. S. Hopkins and J. L. Campbell, *Analyst*, 2015, **140**, 6897–6903.
- 34 C. Liu, J. C. Y. L. Blanc, B. B. Schneider, J. Shields, J. J. Federico, H. Zhang, J. G. Stroh, G. W. Kauffman, D. W. Kung, C. Ieritano, E. Shepherdson, M. Verbuyst, L. Melo, M. Hasan, D. Naser, J. S. Janiszewski, W. S. Hopkins and J. L. Campbell, *ACS Cent. Sci.*, 2017, **3**, 101–109.
- 35 N. J. A. Coughlan, P. J. J. Carr, S. C. Walker, C. Zhou, M. Guna, J. L. Campbell and W. S. Hopkins, *J. Am. Soc. Mass Spectrom.*, 2020, **31**, 405–410.
- 36 A. Haack, C. Polaczek, M. Tsolakis, M. Thinius, H. Kersten and T. Benter, *J. Am. Soc. Mass Spectrom.*, 2020, **31**, 785–795.
- 37 C. Ieritano, D. Rickert, J. Featherstone, J. F. Honek, J. L. Campbell, J. C. Y. L. Blanc, B. B. Schneider and W. S. Hopkins, *J. Am. Soc. Mass Spectrom.*, 2021, **32**, 956–968.
- 38 A. Haack and W. S. Hopkins, *J. Am. Soc. Mass Spectrom.*, 2022, **33**, 2250–2262.
- 39 J. W. L. Lee, M. H. Stockett, E. K. Ashworth, J. E. N. Navarrete, E. Gougoula, D. Garg, M. Ji, B. Zhu, S. Indrajith, H. Zettergren, H. T. Schmidt and J. N. Bull, *J. Chem. Phys.*, 2023, **158**, 174305.
- 40 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman and D. J. Fox, *Gaussian 16 Revision B.01*, Gaussian Inc., Wallingford CT, 2016.
- 41 F. Neese, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2011, **2**, 73–78.
- 42 C. Riplinger, P. Pinski, U. Becker, E. F. Valeev and F. Neese, *J. Chem. Phys.*, 2016, **144**, 024109.
- 43 J.-D. Chai and M. Head-Gordon, *Phys. Chem. Chem. Phys.*, 2008, **10**, 6615–6620.
- 44 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.



- 45 T. M. Henderson, A. F. Izmaylov, G. Scalmani and G. E. Scuseria, *J. Chem. Phys.*, 2009, **131**, 044108.
- 46 K. Raghavachari, G. W. Trucks, J. A. Pople and M. Head-Gordon, *Chem. Phys. Lett.*, 1989, **157**, 479–483.
- 47 A. K. Dutta, M. Saitow, B. Demoulin, F. Neese and R. Izsák, *J. Chem. Phys.*, 2019, **150**, 164123.
- 48 D. G. Liakos and F. Neese, *J. Chem. Theory Comput.*, 2015, **11**, 4054–4063.
- 49 D. G. Liakos, Y. Guo and F. Neese, *J. Phys. Chem. A*, 2019, **124**, 90–100.
- 50 A. Sirohiwal, F. Neese and D. A. Pantazis, *J. Chem. Theory Comput.*, 2021, **17**, 1858–1873.
- 51 S. Gozem and A. I. Krylov, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2021, **12**, e1546.
- 52 W. Zagorec-Marks, M. M. Foreman, J. R. R. Verlet and J. M. Weber, *J. Phys. Chem. Lett.*, 2019, **10**, 7817–7822.
- 53 F. Santoro, A. Lami, R. Improta, J. Bloino and V. Barone, *J. Chem. Phys.*, 2008, **128**, 224311.
- 54 E. Carrascosa, J. N. Bull, M. S. Scholz, N. J. A. Coughlan, S. Olsen, U. Wille and E. J. Bieske, *J. Phys. Chem. Lett.*, 2018, **9**, 2647–2651.
- 55 N. J. A. Coughlan, M. H. Stockett, C. Kjær, E. K. Ashworth, P. C. B. Page, S. R. Meech, S. B. Nielsen, L. Blancafort, W. S. Hopkins and J. N. Bull, *J. Chem. Phys.*, 2021, **155**, 124304.
- 56 B. D. Adamson, N. J. A. Coughlan, R. E. Continetti and E. J. Bieske, *Phys. Chem. Chem. Phys.*, 2013, **15**, 9540–9548.
- 57 J. N. Bull, C. W. West, C. S. Anstöter, G. da Silva, E. J. Bieske and J. R. R. Verlet, *Phys. Chem. Chem. Phys.*, 2019, **21**, 10567–10579.
- 58 J. N. Bull, E. Carrascosa, L. Giacomozzi, E. J. Bieske and M. H. Stockett, *Phys. Chem. Chem. Phys.*, 2018, **20**, 19672–19681.
- 59 J. N. Bull, G. da Silva, M. S. Scholz, E. Carrascosa and E. J. Bieske, *J. Phys. Chem. A*, 2019, **123**, 4419–4430.
- 60 J. N. Bull, J. T. Buntine, E. Carrascosa, M. H. Stockett and E. J. Bieske, *Eur. Phys. J. D*, 2021, **75**, 67.
- 61 J. N. Bull, M. S. Scholz, E. Carrascosa, G. da Silva and E. J. Bieske, *Phys. Rev. Lett.*, 2018, **120**, 223002.

