



Cite this: *Nanoscale*, 2021, **13**, 19341

Encapsulation of lead halide perovskite nanocrystals (NCs) at the single-particle level: strategies and properties

Qixuan Zhong, Muhan Cao * and Qiao Zhang *

Lead halide perovskite NCs (APbX₃, A = formamidinium (FA), methylammonium (MA) or Cs; X = Cl, Br, I or their mixture) have attracted unprecedented attention due to their excellent photophysical properties and wide application prospects. However, the inherent ionic structure of APbX₃ NCs makes them very sensitive to external conditions such as water and oxygen, resulting in poor stability. As a feasible strategy, encapsulation is considered to be effective in improving the stability. In this minireview, we focus on single-particle-level coating, which not only can improve the stability but also maintain the nano effect of the original NCs. This review summarizes the fundamental information on APbX₃ NCs and the necessity of single-particle-level coating. Subsequently, a variety of heterostructures at the single-particle level are introduced in detail. Then, their applications are summarized. Moreover, we discuss the challenges and prospects of the single-particle-level heterostructures based on APbX₃ NCs.

Received 20th August 2021,
Accepted 28th October 2021

DOI: 10.1039/d1nr05478c

rsc.li/nanoscale

1. Introduction

Metal halide perovskites were first discovered in 1893,¹ but it was not until 2009 that they became a focus of attention, owing to their very first application in photovoltaic devices.² In 2015, Kovalenko's group³ reported, for the first time, the preparation of all-inorganic lead halide perovskite CsPbX₃ NCs (X = Cl, Br, I or their mixture). Their excellent photophysical

properties, including high brightness, tunable emission, high color purity (narrow emission peak), high light absorption coefficient, and high defect tolerance, make them promising in optoelectronic fields such as solar cells,⁴ light-emitting diodes (LEDs),^{5,6} and so on. For example, CsPbX₃-based LEDs have achieved a high external quantum efficiency (EQE) of 12.3% for blue LED,⁵ 22% for green LED,⁵ and 23% for red LED.⁷ Considering their outstanding properties and promising performance in devices, APbX₃ NCs have been regarded as one of the next-generation optoelectronics candidates.^{8–11}

For practical applications, stability is a particularly important parameter. APbX₃ NCs are suitable for optoelectronic devices; however, their poor stability is a shortcoming that

Institute of Functional Nano & Soft Materials (FUNSOM), Jiangsu Key Laboratory for Carbon-Based Functional Materials & Devices, Soochow University, No. 199 Ren'ai Road, Suzhou 215123, Jiangsu, People's Republic of China.
E-mail: mhcao@suda.edu.cn, qiaozhang@suda.edu.cn



Qixuan Zhong

Qixuan Zhong received his B.S. degree (2016) from Changshu Institute of Technology. He is now a graduate student in the Institute of Functional Nano & Soft Materials (FUNSOM), Soochow University. His current research interests focus on the synthesis and applications of all-inorganic metal halide perovskite/oxide composite materials.



Muhan Cao

Muhan Cao obtained her Ph. D. degree in Environmental Science and Technology from Hohai University in 2015. Now she is an associate research fellow at Soochow University. Her current research interests focus on the design of metal and semiconductor nanocrystals and their applications in optoelectronic fields.

limits their practical applications. The suboptimal stability of APbX₃ NCs is mainly attributed to their inherent ionic properties as well as the capping ligands that are dynamically balanced on the surface.^{12,13} To enhance durability, a direct and effective method is to prepare highly stable APbX₃ NCs encapsulated in an inert shell, such as polymers and oxides.^{14,15} These shells can significantly improve the stability of APbX₃ NCs against environmental stimuli. Depending on the number of APbX₃ NCs in one shell, the encapsulation strategies can be divided into multi-particle and single-particle coating. Multi-particle coating has been extensively studied because of the simple synthetic methods. The resulting products usually have a large particle size, which may reach up to several micrometers.^{16–22} This large size limits the applications of APbX₃ NCs at the nanometer scale. Moreover, multi-particle aggregations are not conducive to study the influence of the shell on the photoelectric properties of APbX₃ NCs. In contrast, single-particle-level coated APbX₃ NCs, in which one APbX₃ NC is wrapped by one shell, exhibit significantly enhanced stability with the protection of the inert shell. The single-particle-scale modification can well maintain the properties of APbX₃ at the nanoscale, ideal for application in some special fields, such as biological imaging.²³ More importantly, the optical and photoelectric properties can be tuned, enabled by passivation of the NC surface.

In this review article, we focus on summarizing the preparations, properties, and applications of encapsulated APbX₃ NCs at the single-particle level. The first part introduces the design principle for encapsulated APbX₃ NCs. In the second part, APbX₃-based heterostructures at the single-particle level are summarized. The third part presents the applications of APbX₃-based heterostructures. An outlook from our viewpoint is provided to depict the challenges and opportunities with the single-particle encapsulation of APbX₃ NCs in the last part.



Qiao Zhang

Qiao Zhang is a Professor in the Institute of Functional Nano & Soft Materials (FUNSOM), Soochow University, China. He received his Ph.D. from University of California Riverside (2012). After being a postdoctoral Fellow in the Department of Chemistry, University of California Berkeley (2012–2014), he joined Soochow University in 2014. His main research field is molecular-level studies of surface and heterogeneous catalysis,

controllable synthesis of all-inorganic perovskite with tunable optical properties, and their applications.

2. Design principle of encapsulated APbX₃ NCs

2.1 Crystal structure of APbX₃ NCs

Taking CsPbX₃ as an example, Fig. 1a shows a typical cubic CsPbX₃ structure, which is highly symmetric. The Pb ion octahedrally coordinates with six halide (X) ions in a [PbX₆]^{4–} configuration, forming a 3D framework perovskite structure through a continuous array of corner-sharing [PbX₆]^{4–} octahedrons. CsPbX₃ NCs are ionic compounds that can be easily destructed in the presence of polar solvents. The interactions between CsPbX₃ and surface ligands present ionic characteristics as well.^{3,38} The dynamic oleic acid (OA) and oleylamine (OLA) are not tightly bound to the surface of CsPbX₃ NCs.³⁹ Instead, proton exchange between coordinated OLA and OA may induce the detachment of ligands (Fig. 1b), resulting in more trap sites and poor optoelectronic performance.^{39,40} Fully considering the application occasions of CsPbX₃ NCs, environmental stimuli, including light, humidity, heat, and oxygen, will cause the collapse of their structure and quenching of optical properties. Therefore, CsPbX₃ NCs with high stability and excellent photoelectric properties are urgently required.

2.2 Single-particle-level coated APbX₃ NCs

In recent years, some works have been reported to prepare highly stable core/shell APbX₃ NCs. Researchers have developed a variety of strategies for encapsulation of APbX₃ NCs with inert materials to enhance their stability against environmental stimuli.^{13,41–43} The multi-particle coating strategy has been widely reported because of its mild reaction conditions. Amorphous inert materials, including polymers,^{44–46} silica,^{22,47} alumina,²⁰ etc., are usually selected as shell materials to prepare stable structures. Although the coated perovskites exhibit excellent stability, the obtained samples are usually large and even in bulk form, no longer having the superior properties of nanoscale particles that can be used for important applications, e.g., nanoscale optical research and biological fields. Therefore, it is urgently desired but still challenging to synthesize nanosized encapsulated APbX₃ NCs that are normally single-particle coated.

Here, we list the difficulties in preparing single-particle coated APbX₃ NCs: (1) surface ligand intervention. Most of the

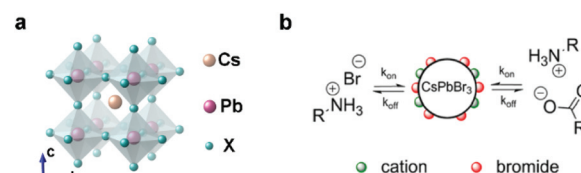


Fig. 1 (a) Schematic of the cubic perovskite lattice. (b) Schematic representation of the dynamic surface stabilization by oleylammonium bromide, oleylammonium oleate and OLA. Reproduced with permission from ref. 39. Copyright 2016 American Chemistry Society.

reported APbX₃ NCs are hydrophobic because their surface is covered by long-alkyl-chain ligands. Since some of the reported shell materials, such as SiO₂ and Al₂O₃, are hydrophilic, it is not easy to grow hydrophilic materials onto the hydrophobic surface directly. (2) Restriction of overgrowth conditions. Some shell materials have to be prepared in the presence of high temperature or polar solvents (such as water). However, unprotected perovskite materials cannot withstand such harsh preparation conditions. (3) Lattice mismatching. To realize the coating of crystalline materials, lattice matching needs to be considered. Due to the particularity of the perovskite structure, it will be better to introduce lattice-matched materials as the shell of APbX₃ NCs. Many efforts have been made to overcome these problems to prepare single-particle-coated APbX₃ NCs. In the next section, we will summarize some typical cases according to different shell materials.

3. APbX₃-based heterostructures

Depending on shell materials, this section is divided into four parts: APbX₃/oxide, APbX₃/semiconductor, APbX₃/metal, and others. The growth mechanism and photophysical properties of different core/shell heterostructures are discussed in this section, and a comparison of the stability before and after encapsulation is listed in Table 1.

3.1 APbX₃/oxide

Inert oxides with high stability, such as SiO₂ and Al₂O₃, are regarded as alternative shell materials for protecting CsPbX₃ NCs due to their weak light absorption and simple preparation strategies.⁴⁸ Most of the reported structures are multi-particle coatings, probably due to the uncontrollable reaction conditions and the destruction of APbX₃ NCs by polar byproducts. Precisely controlling the growth of oxides on the core is particularly important for achieving single-particle-level coating by oxides. This section introduces successful examples of single-particle oxide-coated APbX₃ NCs.

In 2018, our group reported CsPbX₃/SiO₂²⁴ Janus NCs by combining a water-triggered transformation and a sol-gel strategy (Fig. 2a). Cs₄PbX₆ NCs were synthesized by a hot-injection method, then water was added to the mixture of Cs₄PbX₆ NCs and tetramethoxysilane (TMOS) under vigorous stirring. In this process, hydrophobic capping ligands (OA and OLA) on the surface of Cs₄PbBr₆ NCs are removed when CsBr is stripped out from the hexane/water interface. As a result, one side of the NCs contacting water has larger surface energy, so that SiO₂ can nucleate and grow on that side of CsPbX₃ NCs (Fig. 2b). This method has good universality and can be applied in the preparation of a variety of stable Janus structures. As shown in Fig. 2c, CsPbX₃@SiO₂ Janus structures with different halide composites were successfully prepared, which exhibit high brightness and stability. By using different oxide precursors, we successfully prepared different oxide-coated CsPbX₃ Janus structures, including CsPbBr₃/ZrO₂²⁵ (Fig. 2d) and CsPbBr₃/Ta₂O₅ (Fig. 2e), by introducing Zr(OC₄H₉)₄ and

tantalum(v) ethoxide as the precursor, respectively. With the protection of the oxide shell, the obtained Janus nanoparticles exhibit enhanced stability against water and light (Fig. 2f and g). Due to the small size and colloidal stability of the single-particle coating, for the first time, the coated nanoparticles were spin-coated to form a uniform film that also exhibits high stability. This work confirms the feasibility of coating CsPbX₃ on a single-particle scale by effectively removing the ligands on the surface of CsPbX₃ NCs and provides a new direction for subsequent research.

The successful preparation of Janus structures heralds the improved stability of CsPbX₃ NCs, which however requires further improvement due to incomplete protection. We shifted our attention to fully coated CsPbX₃ NCs at a single-particle level. As an amorphous material, SiO₂ has strong plasticity and is easier to fully coat onto most types of nanoparticles.¹⁵ However, SiO₂ is a hydrophilic material that has a large contact angle (θ) with CsPbX₃ NCs, which are hydrophobic. Our group reported a facile one-pot approach to synthesize CsPbBr₃/SiO₂ core/shell NCs.²⁶ The formation process is shown in Fig. 3a. The precursor solution (CsBr, PbBr₂, OA, OLA, and ammonia dissolved in *N,N*-dimethylformamide) is quickly added into ultradry toluene solution containing TMOS, leading to the precipitation of product. At the very beginning of the reaction, the surface energy of NCs is high, and TMOS rapidly hydrolyzes in the presence of water and nucleates on the surface of CsPbBr₃ NCs. It is found that the formation rates, determined by reaction temperature, precursor species, pH value, *etc.*, are critical for the successful preparation of core/shell NCs. Monodisperse, fully coated core/shell CsPbBr₃/SiO₂ NCs can be realized, in which each nanoparticle has only one CsPbBr₃ NC (Fig. 3b). Benefitting from the full protection of the SiO₂ shell, the as-prepared core/shell structure can maintain good stability under very harsh ultrasonication treatment in aqueous solution (Fig. 3c). The full coating allows CsPbBr₃ NCs to be stably dispersed in aqueous solutions, which indicates potential in biological applications such as fluorescence imaging *in vivo*. The disadvantage of this method is that precise control of reaction conditions is required.

In addition to these examples, scientists have devoted many efforts to the preparation of core/shell structured perovskite NCs with different components. Tang *et al.*²⁷ reported CsPbMnX₃/SiO₂ (X = Br, Cl) core/shell NCs with two emission peaks synthesized *via* a facile reverse microemulsion method at room temperature, which can be fabricated into a stable white LED. CsPbX₃/Al₂O₃ structures with tunable thickness were synthesized through a general colloidal atomic layer deposition (c-ALD) method.⁵⁰ Considering that the hydrophobic ligands on CsPbX₃ have poor affinity with hydrophilic oxides, 3-aminopropyl triethoxysilane (APTES) was introduced as a bridging ligand to modify the surface of CsPbX₃, so the binding force between the oxides and CsPbX₃ is enhanced. After that, CsPbX₃/oxide (SiO₂, ZrO₂, Ta₂O₅) core/shell²⁸ structures were prepared, as shown in Fig. 3d. All of the core/shell NCs exhibit high stability against water, UV-light irradiation

Table 1 Summary of the structures, optical properties and stabilities of single-particle-level coated lead halide perovskite NCs

Shell materials	Compound	Structure	PL peak [nm]	PLQY (%)	Stability before coating	Stability after coating	Ref.
Oxides	CsPbBr ₃ /SiO ₂	Janus	517	80	Remnant PL 17.8% after 7 days water treatment; remnant PL 84% after 6 hours light irradiation; storing in ambient for 1 days	Remnant PL 80% after 7 days water treatment; remnant PL 98% after 6 hours light irradiation; storing in ambient for 4 days	24
	CsPbBr ₃ /ZrO ₂	Janus	514	90	Remnant PL 5% after 8 days water treatment; remnant PL 70% after 3 hours light irradiation	Remnant PL 80% after 8 days water treatment; remnant PL 90% after 3 hours light irradiation	25
	CsPbBr ₃ /SiO ₂	Core/shell	501	90	Remnant PL 0% after storing in water for 16 min under ultrasonic treatment; storing in ambient <3 days	Remnant PL 112% after storing in water for 40 min under ultrasonic treatment; storing in ambient for 28 days	26
	CsPbMnX ₃ /SiO ₂	Core/shell	446/607	54.7	Remnant PL 0% after storing in water for 6 days; remnant PL 50% after heating at 100 °C for 1 hour	Remnant PL 90% after storing in water for 6 days; remnant PL 90% after heating at 100 °C for 1 hour	27
	CsPbI ₃ /SiO ₂ CsPbI ₃ /ZrO ₂ CsPbI ₃ /TaO ₅	Core/shell	696	51.1 57.7 59.7	Remnant PL 52–64% after 2 hours water treatment; remnant PL 67.2% after heating at 100 °C; remnant PL 25.8% after 3 hours light irradiation	Remnant PL 95–108% after 2 hours water treatment; remnant PL 85–90% after heating at 100 °C; remnant PL 52.7–62.1% after 3 hours light irradiation	28
	CsPbBr ₃ /TiO ₂	Core/shell	~515	—	Remnant PL 0% after 15 min water treatment; remnant PL 37% after 24 hours light irradiation	Remnant PL 85% after 84 days water treatment; remnant PL 75% after 24 hours light irradiation	29
Semiconductors	CsPbBr ₃ /Ag ₂ S	Core/island	520	82	Remnant PL 0% after 4 days water treatment	Remnant PL 6.35% after 30 days water treatment	30
	CsPbBr ₃ /ZnS	Core/shell	510	—	Remnant PL 0% after 2 hours light irradiation	Remnant PL 60% after 48 hours light irradiation	31
	CsPbBr ₃ /Rb ₄ PbBr ₆	Core/shell	507	85	Remnant PL 0% after 2 hours light irradiation	Remnant PL 90% after 42 hours light irradiation	32
	CsPbBr ₃ /CdS	Core/shell	510	90	Remnant PL 65% after 14 hours light irradiation; remnant PL ~0% after heating at 60 °C for 9 hours	Remnant PL 80% after 14 hours light irradiation; remnant PL 40% after heating at 60 °C for 9 hours	33
Metal	CsPbBr ₃ /Pt single atoms	Core/island	510	0.1	Aggregation after 20 min light irradiation by xenon lamp	Maintain structure after 60 min light irradiation by xenon lamp	34
Others	CsPbBr ₃ /PbSO ₄	Core/shell	510	92	Remnant PL 44% after 150 min ethanol treatment; remnant PL 50% after 90 min acetone treatment; remnant PL 80% after 300 min light irradiation	Remnant PL 75% after 150 min ethanol treatment; remnant PL 90% after 90 min acetone treatment; remnant PL 105% after 300 min light irradiation	35
	CsPbBr ₃ /ZnSO ₄	Core/shell	515	95	Anion-exchange after 40 min	No anion-exchange after 48 h	36
	APbX ₃ @PAA- <i>b</i> -PS (A = FA or Cs)	Core/shell	460–650	30–61	Quenching with added water; remnant PL 18% after 60 min light irradiation	Remnant PL 100% after storing in water for 23 days; remnant PL 92% after 60 min light irradiation	37

and thermal treatment (Fig. 3e), indicating the effective protection from the oxide shells. Surface engineering is considered an effective strategy to achieve single-particle-level full coating.

Dual-shelled MAPbBr₃ NCs coated by polymer and SiO₂ were also successfully prepared.⁴⁹ As shown in Fig. 3f, star-like copolymers, as nanoreactors, were designed to render the growth of core/shell NCs with controlled yet tunable perovskite core diameter, SiO₂ shell thickness and surface chemistry. The colloidal stability, chemical composition stability, photostability, and water stability of MAPbBr₃ NCs are dramatically improved with the protection of dual shells. This strategy can realize controllable single-particle-level coating through a

smart but complex design, and we look forward to the development of a simpler and easier method.

3.2 APbX₃/semiconductor

As inert materials, oxides are widely used as shell materials. One disadvantage of oxides is that most of them are insulators, limiting the charge transfer between particles. In contrast, stable semiconductor materials with good carrier transport performance are good candidates. Zheng and coworkers²⁹ reported CsPbBr₃/TiO₂ core/shell NCs by calcination at 300 °C for 5 h, which allow for efficient photoinduced charge transfer and enhanced photoelectric activity during water treatment.

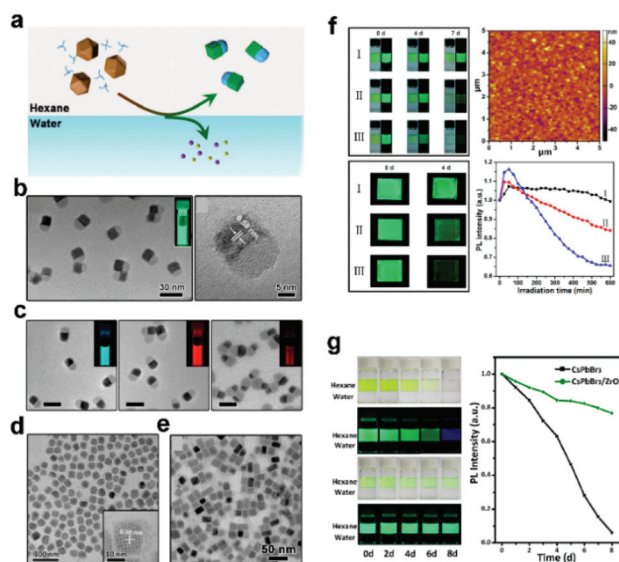


Fig. 2 (a) Schematic illustration of the formation of $\text{CsPbX}_3/\text{oxide}$ Janus NCs. TEM images of (b) $\text{CsPbBr}_3/\text{SiO}_2$, (c) $\text{CsPbX}_3/\text{SiO}_2$ ($X = \text{Cl}/\text{Br}$ or Br/I), (d) $\text{CsPbBr}_3/\text{ZrO}_2$, and (e) $\text{CsPbBr}_3/\text{Ta}_2\text{O}_5$ Janus NCs. Inset shows the corresponding photographs of the colloidal solutions. (f) High stability of $\text{CsPbBr}_3/\text{SiO}_2$ Janus NCs. Photographs show the stability against water and storage in humid air (40 °C and humidity of 75%) of (I) Janus $\text{CsPbBr}_3/\text{SiO}_2$ and CsPbBr_3 obtained from (II) water-triggered and (III) hot-injection methods. Atomic force microscope (AFM) image of $\text{CsPbBr}_3/\text{SiO}_2$ NC thin film. Photostability of different thin films under irradiation of 375 nm UV light. Reproduced with permission from ref. 24. Copyright 2018 American Chemistry Society. (g) High stability of Janus $\text{CsPbBr}_3/\text{ZrO}_2$ NCs. Photographs show the stability against water and the corresponding change of PL intensity. Reproduced with permission from ref. 25. Copyright 2019 American Chemistry Society.

The decoration of CsPbX_3 NCs with semiconductors has attracted widespread interest to improve the stability and modify the photoelectric properties.

In addition to conventional semiconductor materials, chalcogenide quantum dots (QDs) have been widely introduced to modify APbX_3 NCs to regulate their photophysical properties. When semiconductors with different band gaps are compounded, different types of heterostructures can be constructed for different occasions. When applied to perovskites, the biggest challenge here is that the preparation of QDs usually requires high-temperature conditions (often higher than 250 °C) that are not friendly to perovskite NCs. For example, $\text{CsPb}(\text{Br}/\text{I})_3/\text{ZnS}$ heterostructure⁵³ with enhanced stability and tunable photoluminescence was prepared by a solution-phase method, in which ZnS QDs are decorated on the surface of $\text{CsPb}(\text{Br}/\text{I})_3$ NCs. $\text{CsPbX}_3/\text{PbS}$ heterostructures⁵¹ were synthesized *via in situ* growth of PbS using hexamethyldisilathiane as the sulfur precursor. As shown in Fig. 4a and b, $\text{CsPbX}_3/\text{PbS}$ NCs show tunable dual emission feature with the visible and near-infrared (NIR) PL corresponding to CsPbX_3 and PbS, respectively. The fluorescence of the complex can be tuned by varying the size of PbS QDs, which is realized by adjusting the reaction time (Fig. 4c and d). $\text{CsPbBr}_3/\text{PbS}$ is a



Fig. 3 (a) Formation process of $\text{CsPbBr}_3/\text{SiO}_2$ core/shell NCs. (b) TEM and EDS mapping images of $\text{CsPbBr}_3/\text{SiO}_2$ core/shell NCs. (c) High stability of $\text{CsPbBr}_3/\text{SiO}_2$ core/shell NCs against water treatment. Reproduced with permission from ref. 26. Copyright 2018 American Chemistry Society. (d) Synthetic strategy for $\text{CsPbI}_3/\text{SiO}_2$ core/shell NCs using APTES as a linking ligand. (e) High stability of $\text{CsPbI}_3/\text{SiO}_2$ core/shell NCs against water treatment. Reproduced with permission from ref. 28. Copyright 2021 Royal Society of Chemistry. (f) Stepwise representation of the synthetic route to PS-capped $\text{MAPbBr}_3/\text{SiO}_2$ core/shell NCs and PEO-capped $\text{MAPbBr}_3/\text{SiO}_2$ core/shell NCs by exploiting the star-like $\text{P4VP-}b\text{-PtBA-}b\text{-PS}$ and $\text{P4VP-}b\text{-PtBA-}b\text{-PEO}$ as nanoreactors, respectively. CD, cyclodextrin; BMP, 2-bromo-2-methylpropionate; and TOABr, tetraoctylammonium bromide. Reproduced with permission from ref. 49. Copyright 2019 Creative.

typical type I heterojunction. As shown in Fig. 4e and f, these tiny PbS NCs provide channels for the exciton nonradiative recombination, which can quench the PL. Moreover, results prove that exciton energy effectively transfers from CsPbBr_3 to PbS in $\text{CsPbBr}_3/\text{PbS}$ NCs when CsPbBr_3 is excited, which may pave the way toward highly efficient QD photovoltaic and optoelectronic devices. Similarly, $\text{CsPbBrI}_2/\text{PbSe}$ ⁵⁴ and $\text{CsPbBr}_3/\text{Ag}_2\text{S}$ ³⁰ heterojunctions were reported. The QDs on CsPbX_3 NCs make a great contribution to the improvement of stability and the tuning of optical properties.

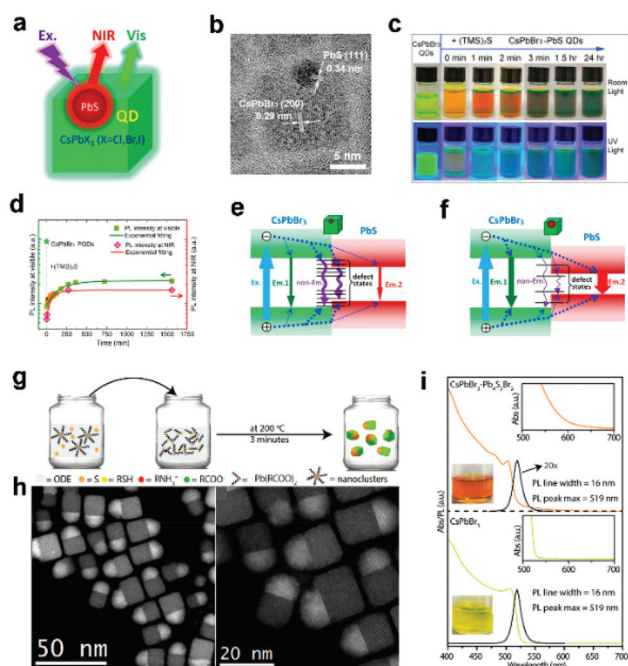


Fig. 4 (a) Schematic illustration of CsPbX₃/PbS heterostructure. (b) TEM image of CsPbBr₃/PbS heterostructure. (c) Photographs of CsPbBr₃ NC solution and CsPbBr₃/PbS NC reaction solution under daylight and UV light. (d) The intensity evolution of PL peaks in visible and NIR regions, respectively. (e, f) Scheme of the band structure and energy transfer channels of the CsPbBr₃/PbS NCs with increasing PbS size, respectively. Reproduced with permission from ref. 51. Copyright 2020 American Chemistry Society. (g) Scheme of the synthesis of CsPbBr₃/Pb₄S₃Br₂ heterostructures. (h) STEM images of CsPbBr₃/Pb₄S₃Br₂ NCs. (i) Absorbance and PL spectra of CsPbBr₃/Pb₄S₃Br₂ heterostructures and CsPbBr₃ NCs. Reproduced with permission from ref. 52. Copyright 2021 American Chemistry Society.

Manna's group⁵² reported an interesting CsPbX₃ Janus heterostructure by epitaxial growth of Pb₄S₃Br₂ onto CsPbBr₃ NCs at high temperature (Fig. 4g). The sublattice of Pb₄S₃Br₂ is similar to that of CsPbBr₃, so that heterostructures can be achieved (Fig. 4h). The overall system is not just the sum of its components but a potentially valuable platform for wave function engineering strategies that are not typically accessible with halide perovskite materials due to compositional alloying. The resulting quasi type II energy level alignment at the heterojunction promotes the rapid splitting of photogenerated excitons, with the photoexcited hole localized in the chalcogenide domain and the electron partially delocalized in the entire nanostructure (Fig. 4i).

Compared to heterostructures obtained from island-like growth, core/shell heterojunctions are highly anticipated. For example, type II CsPbBr₃/ZnS core/shell NCs (Fig. 5a)³¹ were prepared, which exhibit enhanced photoluminescence quantum yield (PLQY) and PL lifetime (Fig. 5b) and improved stability (Fig. 5c). The excellent properties of the core/shell structures lead to their various applications, such as optically pumped LEDs, photodetectors and photocatalysis. Similar CsPbBr₃/CdS^{33,55,56} heterostructures exhibit ultrahigh chemi-

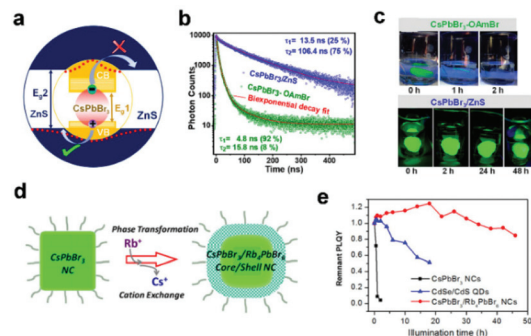


Fig. 5 (a) Schematic showing pseudo type II band alignment at the CsPbBr₃/ZnS core/shell interface, where the electrons are confined inside the core but the holes are delocalized over both core and shell. (b) Comparison of PL decay shows a huge increment in the PL lifetime of CsPbBr₃/ZnS core/shell NCs. (c) Photographs of the films of CsPbBr₃/ZnS core/shell NCs and CsPbBr₃-OLABr NCs dipped in beakers full of water and excited with UV light. Reproduced with permission from ref. 31. Copyright 2020 American Chemistry Society. (d) Schematic diagram of CsPbBr₃/Rb₄PbBr₆ core/shell NCs. (e) Photostability of the solution under illumination with a 450 nm LED light (175 mW cm⁻²). Reproduced with permission from ref. 32. Copyright 2018 American Chemistry Society.

cal stability, nonblinking photoluminescence, and high quantum yield due to the reduced electronic traps within the core/shell structure. To further improve the stability, Liu and coworkers⁵⁷ synthesized a CsPbBr₃/nCdS/Al₂O₃ double-shell structure. The outstanding performance of CsPbBr₃/CdS in optoelectronic applications will be introduced in section 4.

Due to the matching of crystal lattices and similarity of elemental composition, some perovskite-like materials are employed as shell materials to achieve full encapsulation at the single-particle level. Core/shell CsPbX₃/Cs₄PbX₆⁵⁸ structure was first synthesized by a hot-injection method, which opens a new way for the durability enhancement of CsPbX₃ NCs. In work by Tian's group,⁵⁹ amorphous CsPbBr_x is introduced as the shell material to encapsulate CsPbBr₃ NCs, and the resulting blue-emitting CsPbBr₃/CsPbBr_x nanodots display high PLQY and improved stability. Li's group³² prepared CsPbBr₃ NCs enriched with lead bromide on the surface, and then added excess Rb precursor for overgrowth. As a result, CsPbBr₃/Rb₄PbBr₆ NCs are obtained, in which lead bromide is used as an intermediate transition layer to achieve single-particle coating (Fig. 5d). With the excellent passivation of the surface, CsPbBr₃/Rb₄PbBr₆ core/shell NCs show enhanced PLQY and high stability against photo treatment (Fig. 5e), enabling them to be promising in stable optoelectronic devices. CsPbBr₃ can also be introduced as shell material to enhance the properties of FAPbBr₃ NCs.⁶⁰ The disadvantage is that the instability of the ionic structure is not changed.

3.3 CsPbX₃/metal heterostructures

CsPbX₃/metal heterostructures have been carefully investigated, in which photoexcited electron-hole pairs are significantly affected by the introduction of metal. Kamat's group⁶³

reported a CsPbBr₃/Au hybrid architecture in which Au NCs selectively grow on the corners of CsPbBr₃ NCs. This structure is obtained due to facilitating the reduction of AuBr₃ by OLA, followed by Au nucleation on the CsPbBr₃ surface. Moreover, Sheldon *et al.*^{61,64} found that the excess AuBr₃ can be partially reduced to produce nonfluorescent Cs₂Au₂Br₆ NCs. During the process, the addition of PbBr₂ inhibits the substitution of Au for Pb to produce the final CsPbBr₃/Au heterojunction (Fig. 6a–d). The PLQY of CsPbBr₃ NCs is decreased when Au nanoparticles are formed on the corners of CsPbBr₃ NCs, suggesting the charge separation endowed by the introduction of Au. Inspired by these works, our group⁶² demonstrated a reversible fluorescence switching of CsPbI₃ NCs by depositing and de-depositing metallic Ag nanoparticles on CsPbI₃ NCs controlled by light on/off, as shown in Fig. 6e–h. The controllable PL reversibility in the CsPbX₃/Ag composite system favors its application in smart fluorescent devices, such as a rewritable platforms. In our very recent work,³⁴ Pt single atoms were successfully deposited on CsPbBr₃ NCs through a photoassisted approach, in which the surface is partially oxidized first, followed by the anchoring of single Pt atoms through the formation of Pt–O and Pt–Br bonds. The resulting products can be used as efficient and durable catalysts for photocatalytic semi-hydrogenation of propyne. CsPbX₃/metal heterostructures composed of metal and semiconductor characteristics have potential prospects in the fields of optoelectronic devices and catalysis.

3.4 Other heterostructures

In addition to the aforementioned heterojunctions, other core/shell structures have been also studied, including CsPbX₃/PbSO₄,³⁵ CsPbX₃/polymer,⁶⁵ and AuCu/CsPbCl₃.⁶⁶ Our group³⁵ developed a post-treatment method to synthesize a CsPbX₃/PbSO₄ core/shell structure with high PLQY and stability. In this post-treatment method (Fig. 7a), octylammonium ligand is used to increase the affinity of the sulfate with the surface of

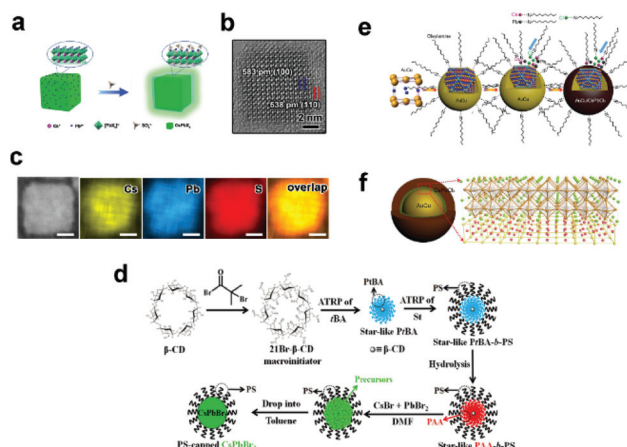


Fig. 7 (a) Schematic illustration of the post-treatment method for synthesizing CsPbBr₃/PbSO₄ core/shell NCs. (b) HRTEM image of one CsPbBr₃/PbSO₄ core/shell NC. (c) HAADF-STEM image and elemental mapping images showing the elemental distribution of Cs, Pb, and S. The scale bars are 3 nm. Reproduced with permission from ref. 35. Copyright 2021 Wiley-VCH. (d) Stepwise representation of crafting hairy all-inorganic perovskite CsPbBr₃ NCs intimately and permanently capped by PS chains via capitalizing on the amphiphilic star-like PAA-*b*-PS diblock copolymer as a nanoreactor. Reproduced with permission from ref. 65. Copyright 2019 Wiley-VCH. (e) Schematic illustration of the AuCu/CsPbCl₃ core/shell NCs' heteroepitaxial growth process. (f) Schematic of the AuCu/CsPbCl₃ core/shell NCs. Reproduced with permission from ref. 66. Copyright 2020 Wiley-VCH.

CsPbX₃ NCs, and then the surface enriched with Pb ions is used to better react with sulfate to produce PbSO₄ shells (Fig. 7b and c). The presence of the PbSO₄ layer favors the protection of the core CsPbX₃ without weakening the charge transfer between CsPbX₃ NCs, making them suitable for stable optoelectronic devices. A general approach was demonstrated for fabricating CsPbX₃/metal sulfate core/shell structure.³⁶ The resulting core/shell NCs with a controlled shell thickness ranging from 1 to 3 nm exhibit enhanced properties in terms of colloidal stability and PLQY, providing a new opportunity for the controllable preparation of core/shell structures.

Lin⁶⁵ reported an amphiphilic star-like block copolymer nanoreactor strategy to prepare CsPbBr₃/polymer NCs with high water and colloidal stabilities (Fig. 7d). The key of this method is the construction of micelles composed of amphiphilic polymers, and the precursor is well dissolved inside the micelles to form CsPbBr₃ NCs. The development of this strategy makes it possible for CsPbX₃ NCs to be used in biological imaging. APbX₃ (A = Cs or FA) NCs encapsulated by amphiphilic polymer (poly(acrylic acid)-*block*-poly(styrene), PAA-*b*-PS) micelles were synthesized by a room-temperature reprecipitation method.³⁷ The as-prepared NCs exhibit high stability against polar solvents and high-flux irradiation. The presence of polymers also prevents ion exchange between NCs, so APbX₃ can be directly used in white LEDs. The single-particle-level encapsulation of perovskites by organic matter is realized by the *in situ* synthesis of stable core-shell structure through

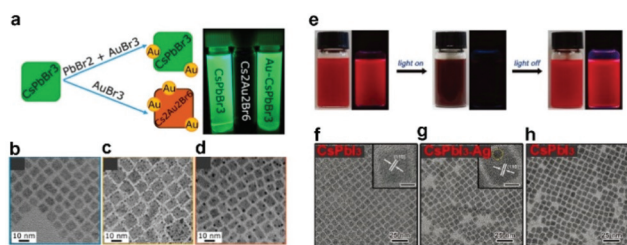


Fig. 6 (a) Schematic illustration of the formation process of CsPbBr₃/Au heterostructure and Cs₂Au₂Br₆ NCs. (b) TEM images of CsPbBr₃ and CsPbBr₃ after gold cation exchange and CsPbBr₃ with gold deposition. Reproduced with permission from ref. 61. Copyright 2017 American Chemistry Society. (e) Photographs showing the color change of the CsPbI₃/AgI composite system induced by light on/off. TEM images of (f) the original CsPbI₃ NCs before irradiation, (g) CsPbI₃/Ag NCs, and (h) CsPbI₃ NCs after being stored in the dark, respectively. The insets in (f) and (g) are the corresponding HRTEM images of CsPbI₃ and CsPbI₃/Ag NCs with a scale bar of 5 nm. Reproduced with permission from ref. 62. Copyright 2019 Royal Society of Chemistry.

the confinement of micelles. This process is more demanding on experimental design and operation requirements.

A typical case of core/shell structure is AuCu/CsPbCl₃ NCs, in which AuCu NC is capped with CsPbCl₃ shell,⁶⁶ as shown in Fig. 7e. The same ligand OLA and hydrophobic property make it easier to overgrow CsPbCl₃ onto the AuCu core. Another significant reason for the core/shell structure is that the CsPbCl₃ shell and AuCu core maintain a high degree of lattice match (Fig. 7f).

4. Applications of core/shell structures

The encapsulation can improve the stability and photoelectric performance of bare CsPbX₃ NCs, making them more suitable for optoelectronic devices, which have high stability and efficiency requirements. For example, owing to the protection of SiO₂ shell and their good film-forming properties (Fig. 3f), CsPbBr₃/SiO₂ Janus NCs²⁴ have been fabricated into white LEDs (WLEDs), which can well maintain white emission after operation for one hour (Fig. 8a). In sharp contrast, the WLED made of bare CsPbBr₃ NCs changed from white to blue within a short period. CsPbX₃/semiconductor heterostructure not only improves the stability and passivates the defects of the perovskite, but also makes it more suitable for photoelectric conversion devices, such as LEDs, solar cells, photodetectors, etc. CsPbI₃/PbS core/shell NCs⁶⁷ were fabricated into electrolu-

minescence LEDs using p-i-n structures due to CsPbI₃ NC films switching from n-type behavior to nearly ambipolar by PbS capping (Fig. 8b). The red LEDs exhibit dramatically enhanced operation stability, and an EQE of 11.8% can be realized. FAPbBr₃/CsPbBr₃ core/shell NCs⁶⁰ with high PLQY and improved stability have been utilized in a green LED, which shows excellent performance with an EQE of 8.1%. Metal-organic framework (MOF)-coated CsPbBr₃ NCs have been demonstrated as the emitter for bright and stable LEDs.⁶⁹ With the protection of MOFs, the as-fabricated LED shows high stability against continuous UV light irradiation, heat, and electrical stress,⁶⁹ and it is free from ion migration, NC merging, and interface corrosion. The operational lifetime of the resulting LEDs has been greatly extended. This work opens up a new way for the large-scale solution-processed fabrication of perovskite-based LEDs with high brightness and high efficiency.

CsPbBr₃ NCs encapsulated in a mesoporous ion-based MOF show significant interfacial charge transfer from CsPbBr₃ NCs to MOF, which is ideal for photocatalysis. The composite exhibits excellent and stable catalytic activities in oxygen reduction reaction (ORR) and oxygen evolution reaction (OER) when working as the synergistic photocathode in the photoassisted Li-O₂ battery. This work not only provides an efficient photoelectrocatalyst for a photoassisted Li-O₂ battery but also highlights the advantage of MOFs in stabilizing perovskites to form synergistic photocatalysts.⁷⁰ When the perovskite is combined with metals, the excited electrons and holes of the perovskite are separated, which is advantageous for catalysis. For example, our group reported the CsPbBr₃/Pt single atom for photocatalytic semi-hydrogenation of propyne.³⁴

With the passivation of PbSe, the CsPbI₃/PbSe film shows a lower trap density and prolonged exciton lifetime, which promote carrier separation and collection.⁶⁸ These improvements result in high-performance CsPbI₃/PbSe NC solar cells with a power conversion efficiency of 13.9% and improved stability against moisture (Fig. 8c). CsPbI₃/PbTe NCs with enhanced structural stability were synthesized by epitaxial growth strategy. The complex was fabricated as stable WLEDs, demonstrating their potential application value in the field of indoor lighting, such as full-visible-spectrum and plant growth LEDs.⁷¹ With the modification of CsPbCl₃ shell, the AuCu/CsPbCl₃ NCs exhibited enhanced light absorption leading to a remarkably enhanced photo-response in the nanohybrid photodetectors, which is more than 30 times that of the counterpart CsPbCl₃ NCs.⁶⁶ This result illustrates the feasibility of implementing the localized surface plasmon resonance light trapping directly in core/shell NCs for high-performance optoelectronics (Fig. 8d). Biocompatible CsPbBr₃/SiO₂ core/shell NCs⁷² with no toxicity, under the protection of SiO₂, were used for bio-imaging and drug delivery. This work paves the way for new biomedical applications and processes. Highly stable CsPbBr₃/phTEOS-TMOS core/shell NCs have been applied to fluorescence and upconversion imaging in living cells, confirming that the perovskite NCs are suitable for various bioimaging applications.²³

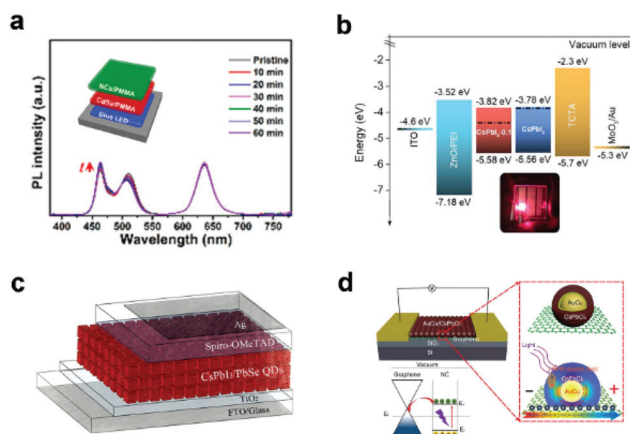


Fig. 8 (a) Time-dependent PL spectra of a WLED under continuous current of 1 mA based on CsPbBr₃/SiO₂ NCs. Inset shows the schematic illustration of the WLED device configuration. Reproduced with permission from ref. 24. Copyright 2018 American Chemistry Society. (b) Device energy level diagram for all functional layers in CsPbI₃/PbS core/shell NC-based electroluminescence LEDs. Inset shows the photograph of the working device. Reproduced with permission from ref. 67. Copyright 2018 American Chemistry Society. (c) Schematics of the device architecture of CsPbI₃/PbSe NC-based solar cells. Reproduced with permission from ref. 68. Copyright 2018 American Chemistry Society. (d) Schematic of the AuCu/CsPbCl₃ NCs nanohybrid photodetector. Reproduced with permission from ref. 66. Copyright 2020 Wiley-VCH.

5. Conclusion and outlook

In this review, we give a brief overview of the fundamental information for APbX₃ NCs and the reasons for their structural instability. The single-particle-level coating structures of various compositions and morphologies are summarized. Based on the excellent performance of the coated APbX₃ NCs, their potential applications in optoelectronic devices such as LEDs, solar cells, photodetectors, and catalysis have been introduced. However, there are still some challenges to be considered in the preparation of single-particle-level structures.

(1) At present, the shell materials and synthetic methods are monotonous. Most of the shells are oxides and chalcogenides, and the synthetic strategies focus on hot-injection and post-treatment. It is necessary to develop more shell materials and preparation approaches that can be used to improve stability and photoelectronic performance.

(2) Unlike the thickness of the shell layer grown on the surface of metal and semiconductor NCs, which can be easily controlled, it is still very difficult to tune the thickness of the shell layer on APbX₃ NCs. The influence of shell thickness on the photoelectric properties of APbX₃ NCs is worthy of attention.

(3) The coating of inert materials shows good stability against water treatment. Due to the insufficient compactness of the shell material or the existence of hydrophobic ligands, it is however impossible to obtain long-term dispersion of core/shell structures in aqueous solution, which limits their practical application.

Conflicts of interest

The authors declare no potential conflicts of interest.

Acknowledgements

This work is supported by the National Natural Science Foundation of China (51922073), Natural Science Foundation of Jiangsu Province (BK20180097), and Postdoctoral Science Foundation of China (2019M661923). We acknowledge the financial support from the 111 Project, Collaborative Innovation Center of Suzhou Nano Science, Technology (NANO-CIC) and the Priority Academic Program Development of Jiangsu Higher Education Institutions (PAPD).

Notes and references

- H. L. Wells, *Z. Anorg. Chem.*, 1893, **3**, 195–210.
- A. Kojima, K. Teshima, Y. Shirai and T. Miyasaka, *J. Am. Chem. Soc.*, 2009, **131**, 6050–6051.
- L. Protesescu, S. Yakunin, M. I. Bodnarchuk, F. Krieg, R. Caputo, C. H. Hendon, R. X. Yang, A. Walsh and M. V. Kovalenko, *Nano Lett.*, 2015, **15**, 3692–3696.
- A. Swarnkar, A. R. Marshall, E. M. Sanehira, B. D. Chernomordik, D. T. Moore, J. A. Christians, T. Chakrabarti and J. M. Luther, *Science*, 2016, **354**, 92–95.
- M. Liu, Q. Wan, H. Wang, F. Carulli, X. Sun, W. Zheng, L. Kong, Q. Zhang, C. Zhang, Q. Zhang, S. Brovelli and L. Li, *Nat. Photonics*, 2021, **15**, 379–385.
- J. Song, J. Li, X. Li, L. Xu, Y. Dong and H. Zeng, *Adv. Mater.*, 2015, **27**, 7162–7167.
- Y. K. Wang, F. Yuan, Y. Dong, J. Y. Li, A. Johnston, B. Chen, M. I. Saidaminov, C. Zhou, X. Zheng, Y. Hou, K. Bertens, H. Ebe, D. Ma, Z. Deng, S. Yuan, R. Chen, L. K. Sagar, J. Liu, J. Fan, P. Li, X. Li, Y. Gao, M. K. Fung, Z. H. Lu, O. M. Bakr, L. S. Liao and E. H. Sargent, *Angew. Chem.*, 2021, **60**, 16164–16170.
- M. Lu, Y. Zhang, S. Wang, J. Guo, W. W. Yu and A. L. Rogach, *Adv. Funct. Mater.*, 2019, **29**, 1902008.
- L. N. Quan, B. P. Rand, R. H. Friend, S. G. Mhaisalkar, T. W. Lee and E. H. Sargent, *Chem. Rev.*, 2019, **119**, 7444–7477.
- J. Shamsi, A. S. Urban, M. Imran, L. De Trizio and L. Manna, *Chem. Rev.*, 2019, **119**, 3296–3348.
- A. Dey, J. Ye, A. De, E. Debroye, S. K. Ha, E. Bladt, A. S. Kshirsagar, Z. Wang, J. Yin, Y. Wang, L. N. Quan, F. Yan, M. Gao, X. Li, J. Shamsi, T. Debnath, M. Cao, M. A. Scheel, S. Kumar, J. A. Steele, M. Gerhard, L. Chouhan, K. Xu, X.-G. Wu, Y. Li, Y. Zhang, A. Dutta, C. Han, I. Vincon, A. L. Rogach, A. Nag, A. Samanta, B. A. Korgel, C.-J. Shih, D. R. Gamelin, D. H. Son, H. Zeng, H. Zhong, H. Sun, H. V. Demir, I. G. Scheblykin, I. Mora-Seró, J. K. Stolarczyk, J. Z. Zhang, J. Feldmann, J. Hofkens, J. M. Luther, J. Pérez-Prieto, L. Li, L. Manna, M. I. Bodnarchuk, M. V. Kovalenko, M. B. J. Roeflaers, N. Pradhan, O. F. Mohammed, O. M. Bakr, P. Yang, P. Müller-Buschbaum, P. V. Kamat, Q. Bao, Q. Zhang, R. Krahne, R. E. Galian, S. D. Stranks, S. Bals, V. Biju, W. A. Tisdale, Y. Yan, R. L. Z. Hoye and L. Polavarapu, *ACS Nano*, 2021, **15**, 10775–10981.
- Q. Sun and W.-J. Yin, *J. Am. Chem. Soc.*, 2017, **139**, 14905–14908.
- Y. Zhou and Y. Zhao, *Energy Environ. Sci.*, 2019, **12**, 1495–1511.
- P. Reiss, M. Protiere and L. Li, *Small*, 2009, **5**, 154–168.
- R. Ghosh Chaudhuri and S. Paria, *Chem. Rev.*, 2012, **112**, 2373–2433.
- H. C. Wang, S. Y. Lin, A. C. Tang, B. P. Singh, H. C. Tong, C. Y. Chen, Y. C. Lee, T. L. Tsai and R. S. Liu, *Angew. Chem., Int. Ed.*, 2016, **55**, 7924–7929.
- D. N. Dirin, L. Protesescu, D. Trummer, I. V. Kochetygov, S. Yakunin, F. Krumeich, N. P. Stadie and M. V. Kovalenko, *Nano Lett.*, 2016, **16**, 5866–5874.
- S. Huang, Z. Li, L. Kong, N. Zhu, A. Shan and L. Li, *J. Am. Chem. Soc.*, 2016, **138**, 5749–5752.
- M. Xing, B. Chen, J. Feng, W. Xu, Y. Bai, Y. Zhou, C. Dong, H. Zhong, J. Zhang and Y. Yin, *Chem*, 2019, **5**, 2195–2214.
- Z. Li, L. Kong, S. Huang and L. Li, *Angew. Chem., Int. Ed.*, 2017, **56**, 8134–8138.

- 21 M. N. An, S. Park, R. Brescia, M. Lutfullin, L. Sinatra, O. M. Bakr, L. De Trizio and L. Manna, *ACS Energy Lett.*, 2021, **6**, 900–907.
- 22 L. Xu, J. Chen, J. Song, J. Li, J. Xue, Y. Dong, B. Cai, Q. Shan, B. Han and H. Zeng, *ACS Appl. Mater. Interfaces*, 2017, **9**, 26556–26564.
- 23 P. M. Talianov, O. O. Peltek, M. Masharin, S. Khubezhov, M. A. Baranov, A. Drabavicius, A. S. Timin, L. E. Zelenkov, A. P. Pushkarev, S. V. Makarov and M. V. Zyuzin, *J. Phys. Chem. Lett.*, 2021, **12**, 8991–8998.
- 24 H. Hu, L. Wu, Y. Tan, Q. Zhong, M. Chen, Y. Qiu, D. Yang, B. Sun, Q. Zhang and Y. Yin, *J. Am. Chem. Soc.*, 2018, **140**, 406–412.
- 25 H. Liu, Y. Tan, M. Cao, H. Hu, L. Wu, X. Yu, L. Wang, B. Sun and Q. Zhang, *ACS Nano*, 2019, **13**, 5366–5374.
- 26 Q. Zhong, M. Cao, H. Hu, D. Yang, M. Chen, P. Li, L. Wu and Q. Zhang, *ACS Nano*, 2018, **12**, 8579–8587.
- 27 X. Tang, W. Chen, Z. Liu, J. Du, Z. Yao, Y. Huang, C. Chen, Z. Yang, T. Shi, W. Hu, Z. Zang, Y. Chen and Y. Leng, *Small*, 2019, **15**, 1900484.
- 28 B. Tang, X. Zhao, L. J. Ruan, C. Qin, A. Shu and Y. Ma, *Nanoscale*, 2021, **13**, 10600–10607.
- 29 Z. J. Li, E. Hofman, J. Li, A. H. Davis, C.-H. Tung, L. Z. Wu and W. Zheng, *Adv. Funct. Mater.*, 2017, **28**, 1704288.
- 30 K. Fu, Y. He, B. Zhang, X. Gao and G. Zou, *J. Electroanal. Chem.*, 2020, **858**, 113835.
- 31 V. K. Ravi, S. Saikia, S. Yadav, V. V. Nawale and A. Nag, *ACS Energy Lett.*, 2020, **5**, 1794–1796.
- 32 B. Wang, C. Zhang, S. Huang, Z. Li, L. Kong, L. Jin, J. Wang, K. Wu and L. Li, *ACS Appl. Mater. Interfaces*, 2018, **10**, 23303–23310.
- 33 X. Tang, J. Yang, S. Li, Z. Liu, Z. Hu, J. Hao, J. Du, Y. Leng, H. Qin, X. Lin, Y. Lin, Y. Tian, M. Zhou and Q. Xiong, *Adv. Sci.*, 2019, **6**, 1900412.
- 34 H. Hu, W. Guan, Y. Xu, X. Wang, L. Wu, M. Chen, Q. Zhong, Y. Xu, Y. Li, T.-K. Sham, X. Zhang, L. Wang, M. Cao and Q. Zhang, *ACS Nano*, 2021, **15**, 13129–13139.
- 35 Q. Zhong, J. Liu, S. Chen, P. Li, J. Chen, W. Guan, Y. Qiu, Y. Xu, M. Cao and Q. Zhang, *Adv. Opt. Mater.*, 2021, **9**, 2001763.
- 36 F. Gao, J. Wu, Y. Zhao, T. Song, Z. Deng, P. Wang, Y. Wang and H. Li, *Nanoscale*, 2021, **13**, 10329–10334.
- 37 M. Imran, B. T. Mai, L. Goldoni, M. Cirignano, H. B. Jalali, F. Di Stasio, T. Pellegrino and L. Manna, *ACS Energy Lett.*, 2021, **6**, 2844–2853.
- 38 C. K. MØller, *Nature*, 1958, **182**, 1436–1436.
- 39 J. De Roo, M. Ibanez, P. Geiregat, G. Nedelcu, W. Walravens, J. Maes, J. C. Martins, I. Van Driessche, M. V. Kovalenko and Z. Hens, *ACS Nano*, 2016, **10**, 2071–2081.
- 40 Q. Zhong, M. Cao, Y. Xu, P. Li, Y. Zhang, H. Hu, D. Yang, Y. Xu, L. Wang, Y. Li, X. Zhang and Q. Zhang, *Nano Lett.*, 2019, **19**, 4151–4157.
- 41 Y. Wei, Z. Cheng and J. Lin, *Chem. Soc. Rev.*, 2019, **48**, 310–350.
- 42 D. Yang, X. Li and H. Zeng, *Adv. Mater. Interfaces*, 2018, **5**, 1701662.
- 43 G. H. Ahmed, J. Yin, O. M. Bakr and O. F. Mohammed, *ACS Energy Lett.*, 2021, **6**, 1340–1357.
- 44 H. Wu, S. Lin, R. Wang, X. You and Y. Chi, *Nanoscale*, 2019, **11**, 5557–5563.
- 45 H. Wu, S. Wang, F. Cao, J. Zhou, Q. Wu, H. Wang, X. Li, L. Yin and X. Yang, *Chem. Mater.*, 2019, **31**, 1936–1940.
- 46 Y. Wei, X. Deng, Z. Xie, X. Cai, S. Liang, P. A. Ma, Z. Hou, Z. Cheng and J. Lin, *Adv. Funct. Mater.*, 2017, **27**, 1703535.
- 47 C. Sun, Y. Zhang, C. Ruan, C. Yin, X. Wang, Y. Wang and W. W. Yu, *Adv. Mater.*, 2016, **28**, 10088–10094.
- 48 Y. Duan, D. Y. Wang and R. D. Costa, *Adv. Funct. Mater.*, 2021, 2104634.
- 49 Y. He, Y. J. Yoon, Y. W. Harn, G. V. Biesold-McGee, S. Liang, C. H. Lin, V. V. Tsukruk, N. Thadhani, Z. Kang and Z. Lin, *Sci. Adv.*, 2019, **5**, eaax4424.
- 50 A. Loiudice, M. Strach, S. Saris, D. Chernyshov and R. Buonsanti, *J. Am. Chem. Soc.*, 2019, **141**, 8254–8263.
- 51 X. Zhang, X. Wu, X. Liu, G. Chen, Y. Wang, J. Bao, X. Xu, X. Liu, Q. Zhang, K. Yu, W. Wei, J. Liu, J. Xu, H. Jiang, P. Wang and X. Wang, *J. Am. Chem. Soc.*, 2020, **142**, 4464–4471.
- 52 M. Imran, L. Peng, A. Pianetti, V. Pinchetti, J. Ramade, J. Zito, F. Di Stasio, J. Buha, S. Toso, J. Song, I. Infante, S. Bals, S. Brovelli and L. Manna, *J. Am. Chem. Soc.*, 2021, **143**, 1435–1446.
- 53 W. Chen, J. Hao, W. Hu, Z. Zang, X. Tang, L. Fang, T. Niu and M. Zhou, *Small*, 2017, **13**, 1604085.
- 54 J. Zhang, X. Liu, P. Jiang, H. Chen, Y. Wang, J. Ma, R. Zhang, F. Yang, M. Wang, J. Zhang and G. Tu, *Nano Energy*, 2019, **66**, 104142.
- 55 X. Tang, J. Yang, S. Li, W. Chen, Z. Hu and J. Qiu, *Front. Chem.*, 2019, **7**, 499.
- 56 J. Shi, W. Ge, J. Zhu, M. Saruyama and T. Teranishi, *ACS Appl. Nano Mater.*, 2020, **3**, 7563–7571.
- 57 X. Liu, X. Zhang, L. Li, J. Xu, S. Yu, X. Gong, J. Zhang and H. Yin, *ACS Appl. Mater. Interfaces*, 2019, **11**, 40923–40931.
- 58 C. Jia, H. Li, X. Meng and H. Li, *Chem. Commun.*, 2018, **54**, 6300–6303.
- 59 S. Wang, C. Bi, J. Yuan, L. Zhang and J. Tian, *ACS Energy Lett.*, 2017, **3**, 245–251.
- 60 C. Zhang, S. Wang, X. Li, M. Yuan, L. Turyanska and X. Yang, *Adv. Funct. Mater.*, 2020, **30**, 1910582.
- 61 B. J. Roman, J. Otto, C. Galik, R. Downing and M. Sheldon, *Nano Lett.*, 2017, **17**, 5561–5566.
- 62 P. Li, D. Yang, Q. Zhong, Y. Zhang, M. Chen, S. Jiang, J. Chen, M. Cao, Q. Zhang and Y. Yin, *Nanoscale*, 2019, **11**, 3193–3199.
- 63 S. K. Balakrishnan and P. V. Kamat, *ACS Energy Lett.*, 2017, **2**, 88–93.
- 64 F. A. Rodriguez Ortiz, B. J. Roman, J. R. Wen, N. Mireles Villegas, D. F. Dacres and M. T. Sheldon, *Nanoscale*, 2019, **11**, 18109–18115.
- 65 Y. J. Yoon, Y. Chang, S. Zhang, M. Zhang, S. Pan, Y. He, C. H. Lin, S. Yu, Y. Chen, Z. Wang, Y. Ding, J. Jung, N. Thadhani, V. V. Tsukruk, Z. Kang and Z. Lin, *Adv. Mater.*, 2019, **31**, 1901602.

- 66 M. Gong, M. Alamri, D. Ewing, S. M. Sadeghi and J. Z. Wu, *Adv. Mater.*, 2020, **32**, 2002163.
- 67 X. Zhang, M. Lu, Y. Zhang, H. Wu, X. Shen, W. Zhang, W. Zheng, V. L. Colvin and W. W. Yu, *ACS Cent. Sci.*, 2018, **4**, 1352–1359.
- 68 S. Wang, C. Bi, A. Portniagin, J. Yuan, J. Ning, X. Xiao, X. Zhang, Y. Y. Li, S. V. Kershaw, J. Tian and A. L. Rogach, *ACS Energy Lett.*, 2020, **5**, 2401–2410.
- 69 H. Tsai, S. Shrestha, R. A. Vilá, W. Huang, C. Liu, C.-H. Hou, H.-H. Huang, X. Wen, M. Li, G. Wiederrecht, Y. Cui, M. Cotlet, X. Zhang, X. Ma and W. Nie, *Nat. Photonics*, 2021, **15**, 843–849.
- 70 G. Y. Qiao, D. Guan, S. Yuan, H. Rao, X. Chen, J. A. Wang, J. S. Qin, J. J. Xu and J. Yu, *J. Am. Chem. Soc.*, 2021, **143**, 14253–14260.
- 71 L. Zhang, M. Zhu, Y. Sun, J. Zhang, M. Zhang, H. Zhang, F. Zhou, J. Qu and J. Song, *Nano Energy*, 2021, **90**, 106506.
- 72 P. Kumar, M. Patel, C. Park, H. Han, B. Jeong, H. Kang, R. Patel, W. G. Koh and C. Park, *J. Mater. Chem. B*, 2020, **8**, 10337–10345.