

Fig. 2 Synthesis of borane-capped metal-dihydrides.

from C_{2v} to C_2 is consistent with the trialkylboranes interacting with the hydride ligands at the Fe-center. Infrared spectroscopy (KBr) supported this formulation with a broad Fe–H–B stretch at $\sim 1839\text{ cm}^{-1}$.⁸ Notably, the energy of this absorption is at significantly lower energy than bis(tris(mercaptoimidazolyl)hydroborato)-Fe(II) species that display a bridging borohydride.⁹ Solution magnetic susceptibility studies (25 °C, THF) establish **1** as high-spin Fe(II) ($\mu_{\text{eff}} = 4.6 \pm 0.2\ \mu_{\text{B}}$).

Single-crystal X-ray diffraction (XRD) experiments confirmed **1** as an iron-dihydride with each hydride capped by pyramidalized trialkylboranes (Fe–B = 2.970 Å; $\sum B_x = 318.6^\circ$;

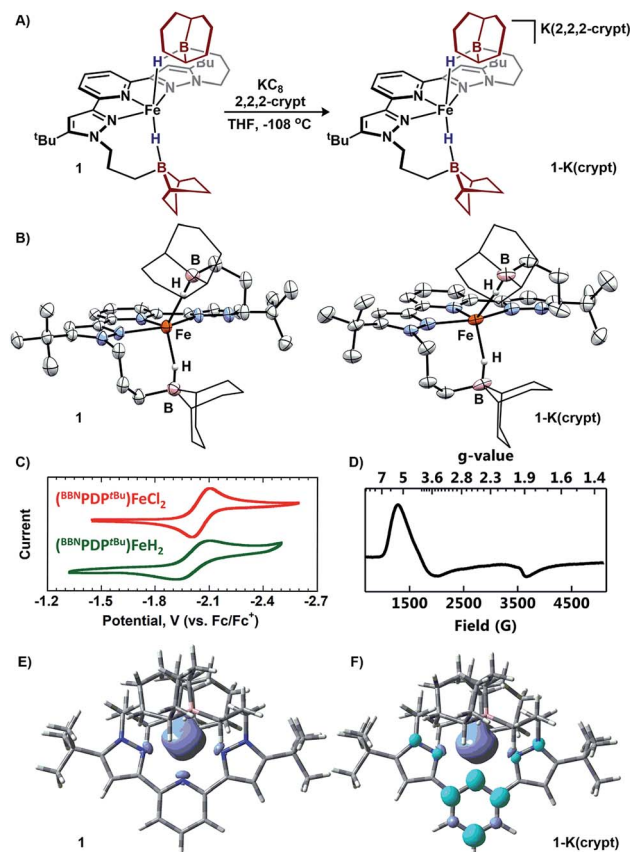


Fig. 3 (A) Reduction of **1**. (B) Molecular structures of **1** and **1-K(crypt)** (50% probability ellipsoids). For clarity, the 9-BBN substituents are displayed in wireframe. (C) Electrochemical comparison between **1** and $(\text{BBNPDP}^{\text{tBu}})\text{FeCl}_2$. (D) X-band EPR of **1-K(crypt)** at 10 K in 1 : 1 toluene/THF. (E and F) Spin-density isosurface plots (0.002 a.u.) of **1** and **1-K(crypt)**, respectively.

Fig. 3B). The geometry about iron is best described as square pyramidal ($\tau_5 = 0.31$). Complex **1** is a rare example of a mono-metallic high-spin Fe(II)-dihydride (or Fe(II)-diborohydride).^{1a,6c,9}

The requirements of the intramolecular acids to stabilize **1** were assessed with a control ligand, $\text{B}^{\text{u}}\text{PDP}^{\text{tBu}}$, where the $-(\text{CH}_2)_3\text{BBN}$ fragments are replaced with *n*-butyl groups (Fig. 2).⁷ Treating $(\text{B}^{\text{u}}\text{PDP}^{\text{tBu}})\text{FeBr}_2$ with 2 equiv. KBHET_3 afforded intractable mixtures, precluding characterization (see ESI[†]). We propose that the appended trialkylboranes in **1** serve two important roles: (1) to stabilize the hydride ligands, and (2) to decrease the entropic penalty for stabilization (intramolecular 9-BBN vs. exogenous BET_3). Attempts to install another donor ligand to **1** and enforce a low-spin configuration were unsuccessful: **1** was unreactive toward stoichiometric 1,4-diazabicyclo[2.2.2]octane, NMe_3 , and PMe_3 , but decomposed when treated with 4-dimethylaminopyridine.

To assess the requirements of the metal center and provide additional spectroscopic characterization, we synthesized the zinc analogue, $(\text{BBNPDP}^{\text{tBu}})\text{ZnH}_2$, from $(\text{BBNPDP}^{\text{tBu}})\text{ZnI}_2$ (Fig. 2). ^{11}B -NMR spectroscopy (THF) of $(\text{BBNPDP}^{\text{tBu}})\text{ZnH}_2$ revealed an upfield resonance at 5.43 ppm, consistent with a tetrahedral boron center.¹⁰ IR spectroscopy (KBr) revealed a broad Zn–H–B stretch at $\sim 1775\text{ cm}^{-1}$ that shifts upon deuterium labeling to $\sim 1300\text{ cm}^{-1}$.¹¹ ^2H -NMR spectroscopy established the assignment of the hydride resonance at 0.65 ppm, which is comparable to other $\kappa^1\text{-Zn}(\text{BH}_4)/\text{Zn}(\text{BH}_3\text{R})$ complexes.¹² The molecular structure of $(\text{BBNPDP}^{\text{tBu}})\text{ZnH}_2$ displays bonding metrics analogous to **1** (see ESI[†]).

Electrochemical investigation of **1** using cyclic voltammetry (0.2 M $[\text{Bu}_4\text{N}][\text{PF}_6]$, THF) revealed a quasi-reversible reductive event at -2.06 V vs. Fc/Fc^+ which is minimally shifted from $(\text{BBNPDP}^{\text{tBu}})\text{FeCl}_2$ ($\Delta = +10\text{ mV}$),¹³ despite their different X-type donors (Fig. 3C).¹⁴ The similar redox potentials suggest that, in analogy to the E_4 state of nitrogenase, reducing equivalents can be delivered at a near constant potential to a Fe(II) state, when stored as bridging hydride equivalents.¹⁵ To prepare the reduced complex, 1 equiv. of KC_8 was added to a freshly thawed THF solution of **1** in the presence of 2,2,2-cryptand, resulting in an immediate color change to a vibrant green species assigned as $[\text{K}(2,2,2\text{-cryptand})][(\text{BBNPDP}^{\text{tBu}})\text{FeH}_2]$ (**1-K(crypt)**; Fig. 3A).¹⁶ Investigation of **1-K(crypt)** by ^1H -NMR spectroscopy (THF) revealed a paramagnetically shifted spectrum with resonances ranging -70.6 – -95.0 ppm with solution C_2 symmetry. Samples of **1-K(crypt)** are less stable than **1** and have an approximate half-life of 12 h in THF at room temperature. Upon reduction, the Fe–H–B infrared absorption shifts to higher energy ($\sim 1866\text{ cm}^{-1}$; KBr). Solution magnetic susceptibility studies (25 °C, THF) of **1-K(crypt)** are consistent with an $S = 3/2$ complex ($\mu_{\text{eff}} = 4.1 \pm 0.1\ \mu_{\text{B}}$).¹⁷ X-band EPR spectroscopy was employed to confirm the spin-state. Regardless of coordination environment or geometry, $S = 1/2$ iron complexes, typically exhibit *g* values near the free electron value,¹⁸ while high-spin complexes exhibit *g*_x-tensors > 3.5 .^{2,19} **1-K(crypt)** displays a broad rhombic signal with *g* values of 5.6, 3.97, and 1.82 at 10 K in 1 : 1 THF/toluene glass (Fig. 3D), which suggests **1-K(crypt)** is best described as high-spin.²⁰ Reduced iron complexes with low coordination numbers (2, 3, occasionally 4) are often high-spin,^{19c,19e,21} while



those stabilized by strong-field ligands (*i.e.* NHC, porphyrin, phosphines) are low-spin. **1-K(crypt)** is a rare example of a high-spin complex with a coordination number ≥ 5 .^{19b}

For structural comparison, **1-K(crypt)** was examined by XRD. Data refinement revealed the anionic portion of **1-K(crypt)** to be geometrically similar ($\tau_5 = 0.36$) to **1** (Fig. 3B). The interacting trialkylboranes are equidistant from the metal center in **1** and **1-K(crypt)**. Upon reduction, the Fe–N_{pyridine} bond distance decreases from 2.178(5) to 2.021(6) consistent with enhanced π -backbonding.²² DFT optimized geometries for the high-spin configuration of **1** and **1-K(crypt)** are in agreement with their crystallographic structures.²³ The spin-density for **1** is localized on iron, consistent with a high-spin Fe(II) description (Fig. 3E). For **1-K(crypt)**, significant spin-density is localized on the pyridyl-moiety of the chelate. The calculated β -SOMO for **1-K(crypt)** corresponding to reduction is primarily comprised of ligand- π^* orbitals on the pyridyl moiety with minimal Fe contribution (19%). This analysis is consistent with reduction of **1** affording an $S = 3/2$ system through antiferromagnetic coupling of high-spin Fe(II) with [^BBNPDP^tBu]¹⁻.²⁴ The calculated electronic structure is consistent with crystallographic bond metrics.²⁵ Pyridyl- π^* population is reflected by C₃–C₄ bond elongation (**1**: 1.377(6); **1-K(crypt)**: 1.433(12)_{ave} Å; Fig. S85[†]).

We sought to examine the generality of the trialkylborane Lewis acids to stabilize both the accumulated hydrides and other small molecules. Addition of E–H substrates (E = NH₂, NHMe, NHPH, OH, PPh, SPh) to freshly thawed THF solutions of **1** afford **2-E** as orange-tan powders with production of H₂ (Fig. 4).²⁶ **2-E** are high-spin Fe(II)-species with solution magnetic susceptibilities ranging $\mu_{\text{eff}} = 4.5\text{--}5.5 \mu_{\text{B}}$ (THF, 25 °C). **2-OH** and **2-PHPh** provide diagnostic infrared handles with sharp $\nu_{\text{(OH)}}$ / $\nu_{\text{(PH)}}$ absorptions observed at 3630 and 2340 cm⁻¹ (KBr), respectively—each consistent with previously reported M–OH–BR₃²⁷ and Fe-phosphides.²⁸ ¹H-NMR spectroscopy revealed **2-NHMe**, **2-NHPH**, **2-OH**, and **2-PHPh** are C₂ symmetric in solution—consistent with Fe–E–B interactions remaining intact in solution. In contrast, **2-SPh** exhibits a solution C_{2v} symmetric spectrum that broadens upon cooling to –80 °C, consistent with reversible B–S binding; likely a consequence of a weaker B–S interaction.²⁹ This observation was supported by DFT analyses. In contrast to the strong B–O interaction in **2-OH** (favored by ~10 kcal per mol per interaction), the calculated B–S interaction is thermodynamically disfavored by ~9 kcal per mol per interaction. In solution, the B–SPh binding equilibria in **2-SPh** were arrested by treating **2-SPh** with two equiv. NH₃. Competitive binding with NH₃ afforded the ammonia-borane species, (^BBNPDP^tBu)Fe(SPh)₂(NH₃)₂ (**3**), highlighting the utility of moderately acidic groups to reversibly interact with small molecule substrates (Fig. 5).

To assess the structural similarities, **2-E** were examined by XRD (Fig. 4). Each displays a pentacoordinate iron best described as distorted square-pyramidal with interacting trialkylboranes. The B-heteroatom distances range from 1.592(2)–2.0504(19) Å following the trend B–OH < B–NHMe < B–NHPH < B–PPh = B–SPh. The same trend is observed for the Fe-heteroatom distance with **2-OH** displaying the shortest bond (1.9812(13) Å) and **2-PHPh**/**2-SPh** displaying the longest bond

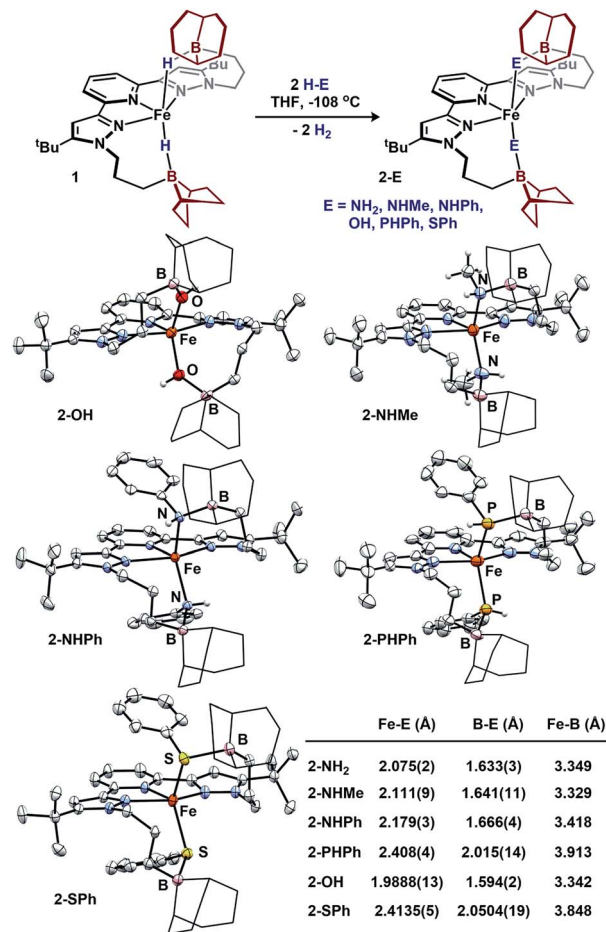


Fig. 4 Formation of **2-E** and molecular structures (50% probability ellipsoids) and averaged bond distances. For clarity, the 9-BBN substituents are displayed in wireframe.

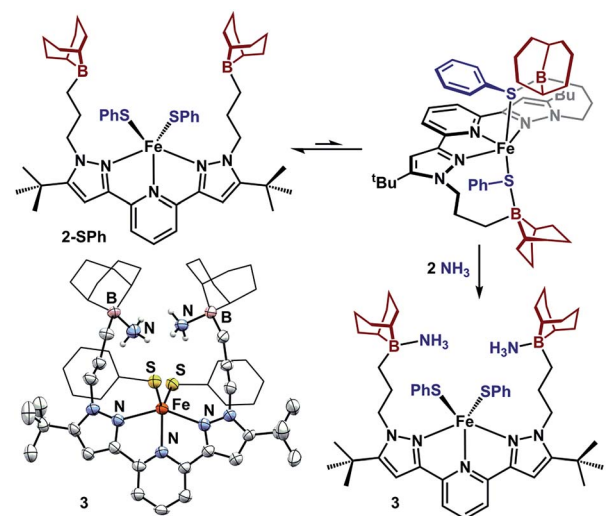


Fig. 5 Reversible acid/base interaction in **2-SPh** and molecular structure of **3** (50% probability ellipsoids). For clarity, the 9-BBN and phenyl substituents are displayed in wireframe.



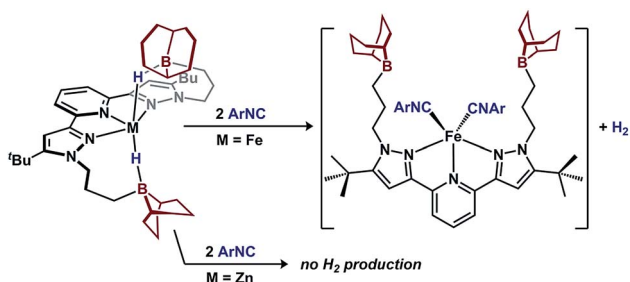


Fig. 6 π -acid induced H_2 extrusion from **1**.

(2.408(4) and 2.4134(5) Å, respectively). The Fe–E bonds are longer than most reported examples, and we attribute this difference to due to quenching the ligand lone-pair by the trialkylborane.^{4d,19b,28,30} Upon substituting the B–S bond of **2-SPH** with NH_3 to afford **3**, the Fe–S bond distance slightly decreases (0.02 Å) while the Fe–S–Ph angle expands (90.06(6) to 110.5(3)_{ave}°). The ability of the appended trialkylborane to stabilize various sizes of small molecules at the metal center is highlighted by comparing the Fe–B distances of **1** and **2-PHPH** which vary by nearly 0.9 Å (2.97 vs. 3.91 Å). The isolation of **2-E** highlights the utility of secondary coordination sphere interactions to not only stabilize hydrides at a metal center, but also to stabilize intermediate species following H_2 loss. Attempts to synthesize most compounds analogous to **2-E** from (^{Bu}PDP^{tBu})FeBr₂ through established methods were unsuccessful (see ESI†).^{19b,28,30b,31} The (^{Bu}PDP^{tBu})Fe(SPh)₂ complex was the only isolable compound, which is consistent with the observed stability of **3** in the absence of Lewis acid/base interactions.

The acid/base interactions can be viewed as an avenue for redox-leveling: the reduction potentials vary by 340 mV across the series of compounds (**1**, **2-E**, and (^{Bu}PDP^{tBu})FeX₂ (X = Cl, Br); Table S3†) despite their disparate ligand donor properties.³² Upon sequestering the Lewis acid of **2-SPH** with NH_3 to form **3**, the redox potential shifts more negative by 80 mV (–2.21 V vs. Fe/Fc⁺). This modest shift is consistent with the weak or dynamic B–S interaction of **2-SPH**. Further description of the Fe–E–B interactions were obtained from their DFT optimized structures.³³ The LUMOs of the β manifold for **2-E** provide insight into their reduction potentials.³⁴ The boron Lewis acids temper the donor properties of the X-type donor ligand, lowering their energy and percent composition in the frontier orbitals (see ESI† for details). The LUMOs of **1** and **2-E** span a modest range of 0.51 eV. In contrast, their computationally investigated truncated (^{Me}PDP^{tBu})Fe(E)₂ counterparts range 1.72 eV. In analogy to the near constant operating potential for N_2 reduction in nitrogenase where $\text{H}^-/\text{N}_x\text{H}_y$ components are present throughout the sequence, the [(^{Bu}PDP^{tBu})Fe] system accommodates small molecule substrates whose donor properties are regulated through Lewis acid/base interactions.

Given the unusual electronic and structural properties of **1**, **1-K(cript)**, and (^{Bu}PDP^{tBu})ZnH₂, Wiberg bond indices for were investigated for each of these complexes to gauge the bonding contribution of the M–H–B interaction (*i.e.* metal–hydride or metal–borohydride character). For each, the interaction is

predominantly borohydride in character with B–H indices ranging 0.65–0.78 and M–H indices ranging 0.14–0.22.³⁵ (^{Bu}PDP^{tBu})ZnH₂ displays the most M–H character whereas **1-K(cript)** displays the most B–H character—consistent with IR spectroscopy. While the data suggest the electronic structures are best described as borohydride, the M–H character varies across the series, indicating M–H reactivity may be accessible.

A key difference between metal–hydrides and borohydrides is the ability of the former to undergo reductive elimination. To test the hydride accumulation hypothesis, *i.e.* whether the bridging hydrides in **1** provide access to Fe(0) by eliminating H_2 , **1** was treated with a π -acid (Fig. 6). Addition of two equiv. 2,4,6-tri-*tert*-butylphenylisocyanide to **1** expels H_2 in 90(2)% yield and affords (^{Bu}PDP^{tBu})Fe(CNAr)₂.³⁶ In contrast, when subjecting the redox-inactive zinc variant, (^{Bu}PDP^{tBu})ZnH₂, to analogous conditions, only 7(3)% H_2 was detected. These results implicate that **1** contains accessible Fe–H character that is available for reactions. In contrast, the reactivity of (^{Bu}PDP^{tBu})ZnH₂ illustrates the inability to extrude H_2 from a borohydride when access to a low-valent metal is not possible.

Conclusions

In summary, we have described a system capable of accumulating hydride equivalents at iron. Intramolecular borane Lewis acid serve a prominent role to not only impart stability to Fe-dihydrides, but also enable previously unobserved electronic structures at iron by regulating the Fe–X-type ligand interactions. Similarities can be drawn between our system and the E₄ state of nitrogenase, where hydrides are accumulated at a near constant potential and addition of small molecule substrates induce H_2 elimination concomitant with substrate stabilization. The geometric flexibility of the appended Lewis acids were previously shown to enable interactions both independent and cooperatively with a metal; here the geometric flexibility is highlighted by their ability to stabilize small molecules of varied sizes. Work is ongoing to interrogate how Lewis acids can be used to further regulate redox transformations at the metal.

Conflicts of interest

There are no conflicts to declare.

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