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Highly swellable pH-responsive nanogels generated by polymerisation-induced self-assembly

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The synthesis of pH-responsive nanogels via reversible addition-fragmentation chain-transfer (RAFT)-mediated polymerisation-induced self-assembly (PISA) in aqueous media, using poly(potassium 3-sulfopropyl methacrylate)₄₄-block-poly(2-hydroxypropyl methacrylate_{1-x%}-*stat*-methacrylic acid_{x%})₃₀₀ (PKSPMA₄₄-*b*-P (HPMA_{1-x%}-*stat*-MAA_{x%})₃₀₀) block copolymers is reported. These nanogels exhibit pronounced and tunable pH-dependent swelling, excellent colloidal stability, and reversibility across multiple pH cycles. These findings advance the design of smart soft materials and establish RAFT-mediated PISA nanogels as promising candidates for pH-triggered delivery systems, biosensing platforms, and adaptive biomaterials.

Introduction

Polymeric nanogels are polymer nanoparticles capable retaining solvent (often referred to as nano-hydrogels in the case of water) and typically exhibit diameters between 20 nm to 200 nm.¹ Nanogels have been broadly investigated and applied in diverse areas, particularly in bioimaging,²⁻⁴ drug-delivery,⁵⁻⁷ and photosensitisation.⁸ Such nanogels can be designed as stimuli-responsive particles responding to pH,⁹⁻¹⁷ temperature,¹⁸⁻²² redox agents,^{23,24} light,^{25,26} and carbon dioxide.²⁷⁻³¹

Nanogels incorporating pH-responsive building-blocks provide swelling behaviour through the protonation and deprotonation of ionisable moieties. These nanoparticles can undergo size variation based on cationic or anionic pH-responsive moieties, such as methacrylic acid (MAA),^{16,17,32} alginate,¹⁴ 2-(methacryloyloxy) ethyl succinate (MES),¹⁵ 2-(diethyl-

amino)ethyl methacrylate (DEAEMA),^{13,33} 2-(dimethylamino) ethyl methacrylate (DMAEMA),^{34,35} and chitosan.^{9,36,37} Anionic monomers such as MAA and alginate typically cause nanogels to expand at high pH owing to deprotonation, while cationic polymers impart the opposite response due to protonation of *e.g.*, amine groups.

Among the various synthetic routes to fabricate polymer nanoparticles, RAFT-mediated PISA has emerged as a robust and versatile approach. PISA generally employs a solvent-philic stabiliser block and a solvent-phobic core-forming block, allowing direct nanoparticle formation at high solids contents.³⁸ Recently, pH-responsive PISA-derived nanogels have been reported. For example, in 2024 Du *et al.* reported a series of pH-responsive nanogels synthesised employing poly(potassium 3-sulfopropyl methacrylate) (PKSPMA) as a macromolecular chain-transfer agent (macro-CTA) and stabiliser block, which was chain-extended with poly(benzyl methacrylate-*stat*-methacrylic acid) as the core-forming block.¹⁵ The swelling behaviour of the resulting nanogels was tuned by adjusting the relative composition of MAA and benzyl methacrylate (BzMA) within the core-forming region. A nanogel with a composition of 20 mol% MAA in the core exhibited an increase in diameter from 51 nm at pH 2 to 84 nm at pH 10. The swelling performance of PISA-derived nanogels was improved in subsequent work by employing a different core-forming monomer.¹⁶ In this case, a relatively hydrophobic carboxylic acid-functional monomer, 2-(methacryloyloxy)ethyl succinate (MES), was used as the core-forming block. The resulting nanogels exhibited more pronounced pH responsiveness, swelling from 66 nm at pH 2 to 143 nm at pH 10, or from 150 nm at pH 2 to 280 nm at pH 10, depending on the formulation. For both formulations the resulting nanogels were investigated as the basis of polymer-nanoparticle-based complex coacervate (PNCC) hydrogels, where the degree of swelling of the selected nanogels directly impacted bulk PNCC gel properties.^{15,17} Thus, the development of PISA-derived nanogel systems with increased swelling properties is of high importance for ongoing work in

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Fig. 2 Hydrodynamic diameter and zeta potential of PKSPMA₄₄-P(HPMA_{1-x%}-MAA_{x%})₃₀₀ plotted against pH. Samples were prepared at 25 °C and 0.10% w/w. (a) S₄₄-H₃₀₀ (b) S₄₄-(H_{95%}-A_{5%})₃₀₀ (c) S₄₄-(H_{90%}-A_{10%})₃₀₀ (d) S₄₄-(H_{85%}-A_{15%})₃₀₀ (e) S₄₄-(H_{80%}-A_{20%})₃₀₀ (f) S₄₄-(H_{75%}-A_{25%})₃₀₀ (g) S₄₄-(H_{70%}-A_{30%})₃₀₀ (h) S₄₄-(isoH_{75%}-A_{25%})₃₀₀.

initial increase in particle diameter is attributed to the deprotonation of carboxylic acid groups in the core of the nanoparticles, leading to enhanced intermolecular electrostatic repulsion, thus swelling of the nanoparticle cores. The zeta-potential of these MAA-containing samples remained highly negative (< -30 mV) across the entire pH range, due to the strongly anionic PKSPMA stabiliser.

A slight reduction in the zeta-potential magnitude was observed upon increasing pH, which can be attributed to the swelling of the nanoparticle core, leading to decreased charge density near the particle surface. Upon further addition of potassium hydroxide during titration, the ionic strength of the colloidal system increased, resulting in charge shielding of the deprotonated carboxylic acid groups and a corresponding slight decrease in particle size observed by DLS measurements.

The swelling behaviour of these nanoparticles is summarised in Table S1 (SI). Interestingly, D_h values at pH 4 decreased moderately with increasing MAA content, falling from 205 nm in the absence of MAA to 123 nm with 30% MAA. This observation may be attributed to the hypothesis that PMAA behaves as a core-compacting component at low pH due to the formation of hydrogen bonds. Conversely, an alternative explanation is that nanoparticles with higher PMAA contents may possess lower aggregation numbers (*i.e.*, fewer chains per particle), possibly due to increased ease of chain exchange, resulting in a larger number of more solvated or loosely associated micelles. These possibilities cannot be confirmed from the current data and will be the subject of future investigation using *e.g.*, small-angle X-ray scattering.

The maximum swelling ratio increased from 1.1 to 2.1 with the inclusion of just 5 mol% MAA in the core forming block and increased further to 4.3 with 25 mol% MAA. This trend is somewhat expected and is attributable to greater amount of core ionisation with increased MAA content, where the

increased electrostatic repulsion between deprotonated chains promotes greater water uptake and thus more pronounced swelling. Nevertheless, no evidence of particle disassociation was observed by DLS or by visible inspection of the resulting dispersions on increasing the pH for samples up to 25 mol% MAA. However, the reduction in swelling degree for S₄₄-(H_{70%}-A_{30%})₃₀₀ may indicate some particle disassociation occurs with high MAA contents in the absence of additional crosslinker, as the core-forming block requires some degree of hydrophobicity to maintain structural integrity. Unfortunately, these observations could not be supported by TEM analysis as these nanogels could not be imaged in the swollen state and will be the subject of additional investigations in the future.

Reversibility of the pH-induced swelling was examined for S₄₄-(H_{75%}-A_{25%})₃₀₀ nanoparticles by adjusting the pH between 4 and 8 repeatedly (Fig. 3). As expected, the nanogels continued to

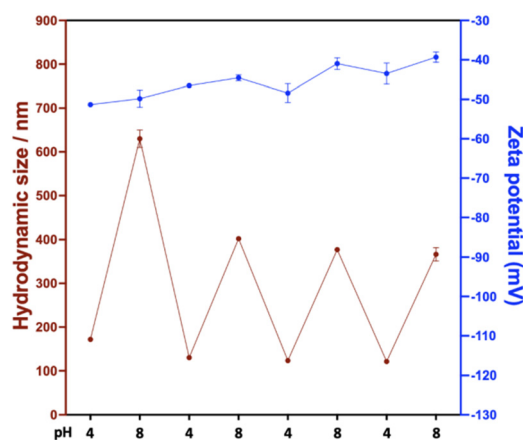


Fig. 3 Variation in particle size with pH cycling between 4 and 8 for S₄₄-(H_{75%}-A_{25%})₃₀₀ dispersions at 25 °C containing 1 mM KCl.



swell/de-swell over at least 4 pH cycles. Interestingly, the swollen diameter measured during the 1st cycle was significantly larger than for subsequent cycles, and the diameters at both pH 4 and pH 8 gradually decreased over successive pH titration cycles. These observations can primarily be attributed to the build-up of ionic strength during these experiments, which screens the charges of the deprotonated carboxylic acid groups within the core and consequently reduces the influx of water into the core of the nanoparticles and thus the degree of swelling. The first cycle shows the most pronounced decrease in diameter because it experiences the largest change in ionic strength. In addition, some segment rearrangement may occur in the absence of cross-linkers. Deprotonated carboxylic acid groups electrostatically repel each other after deprotonation, whereas they likely form hydrogen-bonds in the pristine nanoparticles. This correlates well with the observed decrease in the absolute zeta-potential values recorded during each pH cycle. In addition, the initial dramatic change could also potentially be due to loss of copolymer chains from individual particles in the swollen state. Nevertheless, it is apparent that the majority of the nanogels remain intact, close to their initial unswollen diameter, and continue to display reversible behaviour.

The nanogels discussed thus far employed HPMA monomer which is supplied as a mixture of isomers, with the formulation containing 25 mol% MAA exhibiting the highest degree of swelling. To investigate whether the use of a single HPMA isomer influences nanogel formation and swelling behaviour, 2-hydroxyisopropyl methacrylate (isoHPMA) was used in place of the isomeric mixture for a formulation containing 25 mol% MAA. This isoHPMA-based nanogel displayed a comparable hydrodynamic diameter in the unswollen state but exhibited a markedly greater swelling ratio of 6.2, indicating that the isomeric composition of HPMA plays a role in controlling self-assembly and the resulting nanogel properties. It is somewhat surprising given the very similar structures of HPMA and isoHPMA that such a subtle isomeric difference leads to a markedly increased swelling ratio. This behaviour may arise from increased steric hindrance of the hydroxyl group in isoHPMA, which favours hydrogen bonding between the pendant hydroxyl groups and surrounding water molecules rather than another pendent group. As a result, the isoHPMA-based nanogels are able to take up more water upon deprotonation at higher pH, leading to their enhanced swelling.

PHPMA-based PISA-derived nano-objects are well known to display thermoresponsive behaviour, often undergoing sphere–worm–vesicle transitions, as reported for PGMA–PHPMA systems.^{21,43–45} Such effects are typically observed for formulations prepared using mixed or isoHPMA monomers due to hydrogen bonding between water molecules and the isopropanol hydroxyl groups. Thus, the temperature responsiveness of the S₄₄-(isoH_{75%}-A_{25%})₃₀₀ nanogels was investigated between 1 °C and 45 °C using heating/cooling cycles (Fig. S4, SI). Interestingly, no significant thermal response was observed in the swollen or unswollen state, despite the presence of the isopropyl functionality. This absence of temperature sensitivity is potentially surprising, as one might expect

dissociation of the nanoparticles to unimers on cooling. However, the presence of the carboxylic acid functionality of MAA likely contributes to maintaining structural integrity, preventing temperature-induced disassembly. It appears that the hydrogen bonding occurring predominantly between hydroxyl and carboxylic acid groups within the core-forming region, rather than with water molecules, results in a structure that is relatively insensitive to temperature changes.

Conclusions

In summary, a series of well-defined PKSPMA₄₄-*b*-P(HPMA_{1-x}%-*stat*-MAA_{x%})₃₀₀ amphiphilic block copolymer nanoparticles were successfully synthesised *via* RAFT-mediated aqueous dispersion polymerisation. The polymerisation of KSPMA proceeded in a controlled manner, as confirmed by ¹H NMR and aqueous GPC studies. Subsequent chain extension with HPMA and MAA produced colloiddally stable nanoparticles exhibiting tuneable compositions and pronounced pH-responsive behaviour, as demonstrated by DLS analysis.

Self-assembly occurred through hydrophobic interactions when the MAA content of the 2nd block was 30 mol% or below, without requiring additional cross-linking comonomer. The incorporation of MAA imparted pH-dependent swelling behaviour, with particle diameter increasing upon deprotonation of carboxylic acid groups and reversibly decreasing at low pH. The nanogels also showed reversible size transitions during cyclic pH variation between 4 and 8, demonstrating good structural stability and reproducibility.

Surprisingly, no significant thermal-response was observed for these hydroxypropyl methacrylate-based systems, even when using hydroxyisopropyl methacrylate. This lack of temperature sensitivity is attributed to strong hydrogen bonding between hydroxyl and carboxylic acid groups within the core, forming a compact and thermally stable structure.

Overall, this study demonstrates the synthesis and characterisation of highly swellable, pH-responsive nanogels *via* a RAFT-mediated PISA process. To the best of our knowledge, these nanogels exhibit the most pronounced swelling behaviour of this class of nanogels reported to date. It is worth noting that a reviewer raised the point that the absence of crosslinker may be not that advantageous for certain applications. We acknowledge that the absence of covalent crosslinking may limit mechanical robustness in some contexts; however, the exceptional swelling enabled by this non-crosslinked design provides distinct advantages for applications that rely on pronounced pH-responsive behaviour, offering valuable potential for future applications such as in pH-sensitive drug delivery and other stimuli-responsive soft material systems.

Author contributions

Xueyuan Li: writing – review & editing, writing – original draft, validation, methodology, investigation, formal analysis, data



curation, conceptualisation. Xiaojing Lu: writing – original draft, methodology, investigation, formal analysis. Jian Tang: investigation. Francesca Patel-Burrows: investigation. Lee A. Fielding: writing – review & editing, supervision, resources, project administration, methodology, funding acquisition, formal analysis, data curation, conceptualisation.

Conflicts of interest

There are no conflicts to declare.

Data availability

The data supporting this article have been included as part of the supplementary information (SI). Supplementary information: Materials and methods, GPC data, ¹H NMR spectra, TEM images, particle diameter *versus* temperature, and summary table of swelling behaviour of reported copolymers. See DOI: <https://doi.org/10.1039/d5py01009h>.

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