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Interconnected NiCo₂O₄ nanosheet arrays grown on carbon cloth as a host, adsorber and catalyst for sulfur species enabling high-performance Li–S batteries†

Junli Zhou,^a Xiaolan Yang,^a Yajun Zhang,^a Jinzhu Jia,^a Xinjian He,^a Lin Yu,^{ID} ^{*a}
Yuede Pan,^{ID} ^{*abc} Jingwen Liao,^{ID} ^d Ming Sun^{ID} ^a and Jun He^{ID} ^a

Li–S batteries are a promising next-generation electrochemical energy-storage system due to their high energy density, as well as the abundance and low cost of sulfur. However, the low conductivity of sulfur and Li₂S/Li₂S₂, as well as the dissolution and shuttling of intermediate lithium polysulfides, is a great challenge for high-performance Li–S batteries. Herein, interconnected NiCo₂O₄ nanosheet arrays grown on carbon cloth (CC) are applied as the cathode (S/NiCo₂O₄/CC) in Li–S batteries for accommodating sulfur. The obtained cathode shows high conductivity, high dispersion of sulfur species and excellent polysulfide adsorption and catalytic properties. As a result, significantly higher specific capacity (1480 vs. 1048 mA h g⁻¹ at 0.1C) and greatly enhanced rate performance (624 vs. 215 mA h g⁻¹ at 2C) are obtained for the S/NiCo₂O₄/CC cathode in comparison to S/CC. Further, the S/NiCo₂O₄/CC cathode demonstrates a low capacity decay of 0.060% per cycle over 400 cycles at 0.5C.

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1. Introduction

Rechargeable lithium-ion batteries nowadays are dominating the energy market ranging from electronics to electric vehicles. The main advantage of lithium-ion batteries compared to other batteries such as lead acid and Ni–MH is their high energy density.^{1–5} Further improvement of the energy density can be achieved by using next-generation batteries.^{6–10} The use of batteries with alkaline metals as the anode is a feasible approach for this purpose.^{11–13} Li–S batteries, with high theoretical specific capacity (1675 mA h g⁻¹), high theoretical gravimetric energy density (2600 W h kg⁻¹) and volumetric energy density (2800 W h L⁻¹), remain an attractive choice.^{14–16} The abundance and low cost of sulfur, in contrast to the limited reserves and high price of cobalt and nickel, add to the superiority of Li–S batteries. However, there are several severe

problems for Li–S batteries. Firstly, the low electronic conductivity of sulfur (5×10^{-30} S cm⁻¹) and Li₂S/Li₂S₂ inevitably causes the low utilization of the sulfur cathode.^{17,18} Secondly, there is a significant volume change upon cycling caused due to the density difference between sulfur (2.36 g cm⁻³) and Li₂S (1.66 g cm⁻³),^{19,20} which might lead to structure collapse of the electrode. Thirdly, the high solubility of intermediate lithium polysulfides (LiPS) produced upon cycling causes active material loss and shuttling problems. In respect of the three major issues of electrical insulation, volume change, and polysulfide loss, porous carbon materials physically confining sulfur are proposed.^{21–23} However, due to the weak physical interactions between non-polar carbon materials and polar LiPS it is difficult to effectively inhibit the shuttle effect.^{24–26}

In order to further improve the battery performance, M–X (M = transition metal, X = O, S, N, P, C, etc.)^{27–31} has been widely used as a sulfur host for promoting the kinetics of polysulfide conversion reactions through a strong polar interaction, mitigating polysulfide shuttling and thereby enhancing the battery cycling stability. Recently, spinels (AB₂O₄) emerged as an interesting material hosting sulfur for Li–S batteries. Mixed valences of A (A²⁺/A³⁺) and B (B²⁺/B³⁺) elements provide an activity center for the catalytic conversion of sulfur species.^{32–35}

Herein, carbon cloth (CC) decorated with interconnected NiCo₂O₄ nanosheet arrays as an efficient integrated sulfur host is proposed. The carbon-based current collector ensures excellent electronic conductivity for improved electrochemical kinetics and long-term stability.^{36–38} The high polarity and larger

^aKey Laboratory of Clean Chemistry Technology of Guangdong Higher Education Institutions, Guangzhou Key Laboratory of Clean Transportation Energy Chemistry, Faculty of Chemical Engineering and Light Industry, Guangdong University of Technology, Guangzhou 510006, Guangdong, China. E-mail: gych@gdut.edu.cn

^bInstitute of Energy Innovation, College of Materials Science and Engineering, Taiyuan University of Technology, Taiyuan 030024, China. E-mail: panyuede@tyut.edu.cn

^cKey Laboratory of Advanced Energy Materials Chemistry (Ministry of Education), College of Chemistry, Nankai University, Tianjin 300071, China

^dChinese Academy of Sciences, Guangzhou Institute of Advanced Technology, Guangzhou 511458, Guangdong, China

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surface area increased by introducing the interconnected NiCo₂O₄ nanosheet arrays on CC not only cause homogeneous dispersion of sulfur species but also accommodate the volume changes of sulfur. Moreover, the NiCo₂O₄ nanosheet arrays act as polar host for sulfur, as well as a catalyst to accelerate the charge-transfer process and improve the electrochemical kinetics for the polysulfide conversion, thus greatly retaining the sulfur species in the cathode. Therefore, benefiting from the CC and NiCo₂O₄ nanosheet arrays, the 3D integrated sulfur host greatly improves the conductivity and effectively mitigates polysulfide dissolution. Remarkably, the obtained integrated S/NiCo₂O₄/CC cathode shows significantly higher specific capacity and greatly enhanced rate performance for advanced lithium-sulfur batteries.

2. Results and discussion

Fig. 1a illustrates the two-step preparation procedure of the S/NiCo₂O₄/CC composite. Carbon cloth, the substrate, is composed of carbon fibers with a smooth surface and a diameter of $\sim 9 \mu\text{m}$ (Fig. 1a, b and S1†), which provide an effective platform for the growth of NiCo₂O₄ nanosheets. A rough surface might make the growth of nanosheets, particularly their interconnection, difficult. Similarly, a too small diameter might render the formation of arrays impossible. Through hydrothermal treatment of CC combined with annealing, interconnected NiCo₂O₄ nanosheet arrays were grown on the surface of the carbon fibers of CC, with a coating layer thickness of $\sim 0.5 \mu\text{m}$ (Fig. 1c). In the hydrothermal reaction, the co-precipitation of Ni²⁺, Co²⁺ and OH⁻ (originated from hexamethylenetetramine), as well as the decomposition of CH₃OH, produced the precursor, which has a composition similar to that of (Ni_xCo_y)(OH)_a(CO₃)_b(H₂O)_c and a morphology resembling that of the NiCo₂O₄ nanosheets (Fig. S2a and b†). During annealing, the

precursor was turned into NiCo₂O₄ through the decomposition reaction, releasing CO₂ and H₂O to form NiCo₂O₄.^{39,40} The as-obtained NiCo₂O₄/CC was then immersed in S/CCl₄ solution to allow the infiltration of sulfur. The loading of sulfur could be adjusted by controlling the concentration of the S/CCl₄ solution and the immersion time. Further, melt diffusion was conducted to boost sulfur infiltration on the surface and into the pores of the interconnected NiCo₂O₄ nanosheet arrays.^{41,42} As can be seen from Fig. 1d, sulfur is homogeneously dispersed on the NiCo₂O₄/CC surface. In contrast, sulfur aggregates exist on the surface of the CC fiber for the S/CC composite (Fig. S3†), which was prepared following a similar procedure as described before. Hence, carbon cloth fiber surfaces decorated with NiCo₂O₄ nanosheet arrays show greater affinity for sulfur loading, as compared to the smooth surface of bare CC. This improvement is ascribed to the significantly enlarged surface area and greater sulfur-philic properties enabled by the NiCo₂O₄ nanosheets (Fig. 2e and f).

A detailed examination of the interconnected NiCo₂O₄ nanosheet arrays is shown in Fig. 2. The surface of the carbon cloth fibers is completely covered by vertically aligned interconnected NiCo₂O₄ nanosheets, with no exposure of the original smooth surface of the carbon fibers (Fig. 2a). The NiCo₂O₄ nanosheets form a honeycomb-like structure. In a typical polygon constructed from the NiCo₂O₄ nanosheets, the average side length is 500–1000 nm (Fig. 2b). With even higher magnification, the NiCo₂O₄ nanosheets show a certain degree of transmittance, indicating the thinness of the nanosheets (Fig. 2c). The interconnected NiCo₂O₄ nanosheet arrays significantly improve the surface area, as will be stated below, and provide a platform for sulfur infiltration. The XRD peaks of NiCo₂O₄/CC are assigned to the cubic spinel NiCo₂O₄ phase (JCPDS card no. 20-0781) and the carbon cloth (Fig. 2d). The pore size distribution curve of NiCo₂O₄/CC indicates that the



Fig. 1 Growth of interconnected NiCo₂O₄ nanosheet arrays on CC fibers and morphological characterization. (a) Schematic illustration of the preparation of NiCo₂O₄/CC and S/NiCo₂O₄/CC. Bare CC was treated in an autoclave with Ni²⁺, Co²⁺ and HMTA in methanol at 180 °C before being annealed at 350 °C to obtain NiCo₂O₄/CC. The as-prepared NiCo₂O₄/CC was then immersed in S/CCl₄ solution to load sulfur and treated at 155 °C for melt diffusion of the infiltrated sulfur. SEM images of (b) CC (magnification: 3500), (c) NiCo₂O₄/CC (magnification: 6000) and (d) S/NiCo₂O₄/CC (magnification: 10 000).





Fig. 2 Characterization of NiCo₂O₄/CC. SEM images of the NiCo₂O₄ nanosheets with magnifications of (a) 9000, (b) 20 000 and (c) 70 000, respectively. (d) XRD patterns of NiCo₂O₄/CC and CC. (e) Pore size distribution of NiCo₂O₄/CC and CC. (f) Surface area and pore volume of NiCo₂O₄/CC and CC.

majority of the pores are at around 4 nm, and simultaneously there are micropores and macropores existing (Fig. 2e). Owing to the NiCo₂O₄ nanosheet coating on the carbon fibers, both the surface area and the pore volume are improved two orders of magnitude, from 0.63 to 110.16 $m^2 g^{-1}$ and from 0.003 to 1.21 $cm^3 g^{-1}$, respectively (Fig. 2f). The low surface area of CC is verified by the high smoothness of both the surface (Fig. 1b) and the interior region (Fig. 3b).

Following the characterization of NiCo₂O₄/CC (Fig. 2), S/NiCo₂O₄/CC was carefully examined (Fig. 3). A bird view of the S/NiCo₂O₄/CC fiber reveals a coarse surface and a smooth cross

section (Fig. 3a). The side view of the S/NiCo₂O₄/CC fiber shows a smooth heart-shaped cross section, with a core-shelled fiber structure of S/NiCo₂O₄/CC possessing the core of the CC fiber and the shell of S/NiCo₂O₄ (Fig. 3b). Elemental analyses of S, Ni, Co, O, and C (Fig. 3c and d) further verify the core-shell structure of S/NiCo₂O₄/CC. Sulfur does not exist inside the carbon fiber but is spread all over the coating layer of S/NiCo₂O₄ (Fig. 3c), indicating that the carbon fiber acts only as a current collector and the interconnected NiCo₂O₄ nanosheets as the host for sulfur. In other words, sulfur can stay only in the NiCo₂O₄ layer and will not be able to penetrate into the interior of



Fig. 3 Characterization of S/NiCo₂O₄/CC. (a) SEM images of S/NiCo₂O₄/CC. (b) Cross-sectional view of S/NiCo₂O₄/CC. (c and d) EDX mapping of the cross section of S/NiCo₂O₄/CC with elements S, Ni, Co, O and C. (e) N₂ adsorption-desorption isotherms and (inset) pore size distribution of S/NiCo₂O₄/CC. (f) XRD patterns of S/NiCo₂O₄/CC with different sulfur loadings and sulfur.



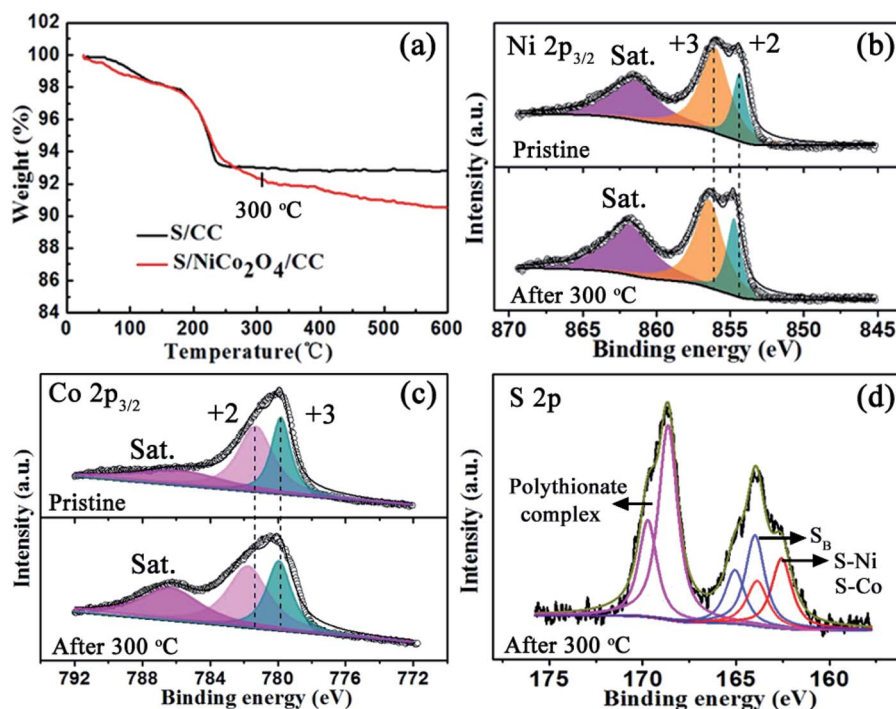


Fig. 4 TGA and XPS characterization of S/NiCo₂O₄/CC. (a) TGA curves of S/CC and S/NiCo₂O₄/CC. High-resolution XPS spectra of (b) Ni 2p_{3/2} and (c) Co 2p_{3/2} for S/NiCo₂O₄/CC before and after heating to 300 °C and (d) S 2p of S/NiCo₂O₄/CC after heating to 300 °C.

the carbon fiber. This observation agrees well with the BET results of bare CC and NiCo₂O₄/CC. While the specific surface area of CC is merely 0.6 m² g⁻¹, with the growth of the nano-sheets the surface area of NiCo₂O₄/CC increases to 110 m² g⁻¹; further, with the loading of sulfur, the surface area drops back to 1.2 m² g⁻¹ (Fig. 3e and S4[†]). As for the other elements Ni, Co, O and C, their dispersion patterns are easy to understand. Ni and Co coexist outside of the carbon fiber; the intensity of O is higher for the exterior than for the interior of the carbon fiber; carbon is only present in the area of the carbon cloth fiber. S/NiCo₂O₄/CC samples with different sulfur loadings (from 1.1 to 3.1 mg cm⁻²) are characterized with XRD, verifying the presence of sulfur with the same orthorhombic structure as pure sulfur powder (Fig. 3f). With the increase of the sulfur content, the intensity of the characteristic peaks for sulfur increases accordingly.

The properties of S/NiCo₂O₄/CC are further explored with TGA and XPS. After heating to 300 °C, the weight of S/CC stabilizes up to 600 °C. The interaction between sulfur and carbon is weak, and therefore sulfur is completely lost below 300 °C, in accordance with the thermodynamic behavior of pure sulfur.^{43,44} In contrast, the weight of S/NiCo₂O₄/CC continues to decrease above 300 °C, which might be ascribed to the strong bonding between elemental sulfur and NiCo₂O₄, as will be shown below with XPS results. The Ni 2p_{3/2} XPS spectrum of the initial S/NiCo₂O₄/CC exhibits a spin-orbit doublet at 856.0 and 854.4 eV, corresponding to Ni³⁺ and Ni²⁺, respectively. Above 300 °C, the peaks for Ni³⁺ and Ni²⁺ shift to 856.5 and 854.7 eV, respectively. Similarly, the two peaks in the Co 2p_{3/2} spectrum shift from 781.3 to 781.8 and from 779.6 eV to 779.9 eV,

respectively, after heating to 300 °C. The peak position variations of Ni 2p_{3/2} and Co 2p_{3/2} spectra might be attributed to the electron transfer between sulfur and Ni/Co atoms, manifesting the strong chemical interaction between NiCo₂O₄ and S₈.^{33,45,46} This speculation is confirmed from the S 2p XPS spectrum (Fig. 4d). For S/NiCo₂O₄/CC above 300 °C, there are two multiple S 2p spectra, which are separated into three sulfur environments. The peaks at 165.0 and 163.9 eV are ascribed to the bridging sulfur (S_B) of S₈. The peaks at 169.7 and 168.6 eV correspond to the polythionate complex, which was produced from the oxidation of sulfur by NiCo₂O₄. Moreover, the peaks for Co²⁺/Co³⁺ and Ni²⁺/Ni³⁺ shift to a higher binding energy by 0.3–0.5 eV above 300 °C, which is ascribed to the strong chemical affinity between sulfur and Ni/Co cations, evidenced by the S 2p peaks at 163.8 and 162.6 eV for S–Ni and S–Co bonding. In contrast, for S/NiCo₂O₄/CC before heating, the interactions of S–Ni and S–Co are evidently weaker (Fig. S5[†]).

As stated above, the S/NiCo₂O₄/CC cathode shows several advantages compared to S/CC. First, because of the two-orders-of-magnitude higher specific surface area of NiCo₂O₄/CC than CC, sulfur dispersion in S/NiCo₂O₄/CC presents significantly higher homogeneity, which might improve the specific capacity, rate performance and cycle life. Second, S–Ni and S–Co interactions render strong bonding between sulfur species and NiCo₂O₄/CC, and hence tend to suppress polysulfide dissolution and diffusion. Third, the interconnected NiCo₂O₄ nano-sheet arrays with large spaces among the sheets allow better electrolyte permeation and solve the volume expansion problem caused by S/Li₂S transition.



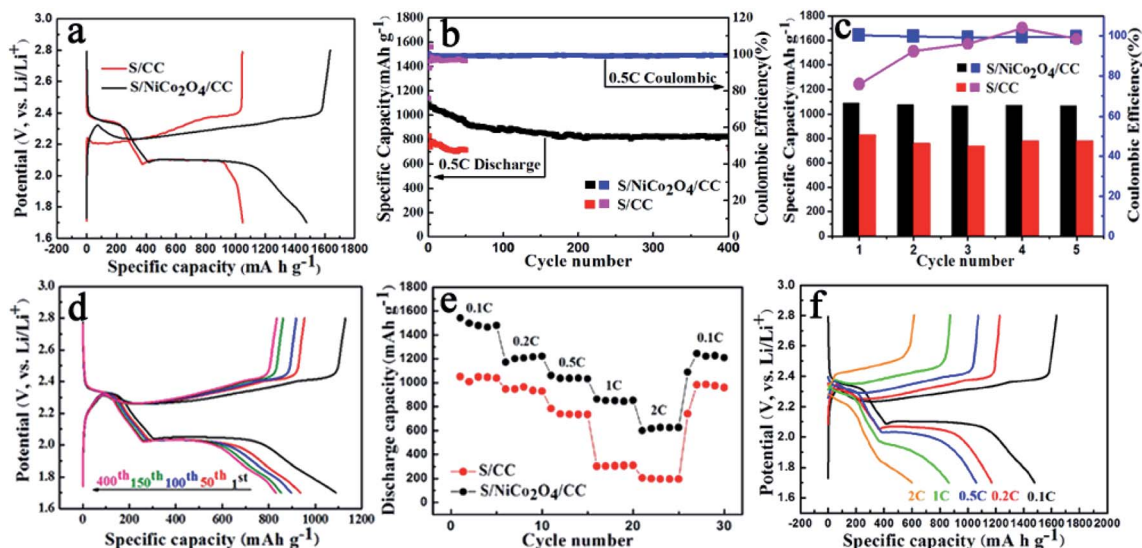


Fig. 5 Electrochemical performance of S/CC and S/NiCo₂O₄/CC (sulfur loading: 1.1–1.3 mg cm⁻²). (a) Comparison of the first charge–discharge profiles at 0.1C for S/CC and S/NiCo₂O₄/CC. (b) Cycling performances and coulombic efficiencies of S/CC and S/NiCo₂O₄/CC at 0.5C. (c) Specific capacities and coulombic efficiencies of S/CC and S/NiCo₂O₄/CC for the first five cycles at 0.5C. (d) Charge–discharge profiles during cycling at 0.5C for S/NiCo₂O₄/CC at the 1st, 50th, 100th, 150th and 400 cycles. (e) Capacities at various rates from 0.1 to 2C. (f) Charge–discharge profiles for S/NiCo₂O₄/CC at various rates from 0.1 to 2C.

In light of the merits mentioned above, S/NiCo₂O₄/CC shows significantly improved performance compared to S/CC (Fig. 5). The S/NiCo₂O₄/CC cathode delivers an initial discharge capacity of 1480 mA h g⁻¹ (~88.4% of the theoretical specific capacity), much higher than the 1048 mA h g⁻¹ for S/CC, demonstrating the functionality and effectiveness of the NiCo₂O₄ nanosheets grown on carbon fibers. This high capacity of 1480 mA h g⁻¹ for the S/NiCo₂O₄/CC cathode is among the best results compared to the reports in the literature.^{47–49} Both the S/NiCo₂O₄/CC and S/CC electrodes show typical discharge and charge curves of a Li–S battery with a lithium anode and an elemental sulfur cathode. There are two discharge plateaus at 2.35 and 2.05 V (vs. Li⁺/Li), which, in consequence, are attributed to the reduction of sulfur (S₈) to soluble long-chain LiPS (Li₂S_x, 4 ≤ x ≤ 8) and then the insoluble short-chain LiPS (Li₂S₂/Li₂S). The electrochemical reactions during the charge process mainly occur between 2.2 and 2.4 V. When increasing the sulfur loading, the capacity drops. At a loading of 3.08 mg cm⁻², the capacity

remains at around 689 mA h g⁻¹, still an acceptable performance for Li–S batteries (Fig. S6[†]). At a higher rate of 0.5C, the S/NiCo₂O₄/CC cathode delivers an initial discharge capacity of 1090 mA h g⁻¹ and a high reversible capacity of 826 mA h g⁻¹ after 400 cycles, corresponding to a high capacity retention of 76% and a low capacity decay rate of mere 0.060% per cycle (Fig. 5b). In contrast, the specific capacities for S/CC are evidently lower than those for S/NiCo₂O₄/CC. For the first five cycles, as clearly shown in Fig. 5c, the capacities for S/CC are much lower. Besides, the initial coulombic efficiency for S/CC is extremely low (<80%), in contrast to the greatly higher coulombic efficiency (>99%) for S/NiCo₂O₄/CC (Fig. 5c and S7[†]). During the 400 cycles at 0.5C, the charge/discharge voltage plateaus remain almost the same (Fig. 5d) and the morphology can be well maintained (Fig. S8[†]), indicating the high stability of the S/NiCo₂O₄/CC cathode. Besides, when further improving the sulfur loading, the S/NiCo₂O₄/CC cathode also showed good cycling stability (Fig. S9[†]).

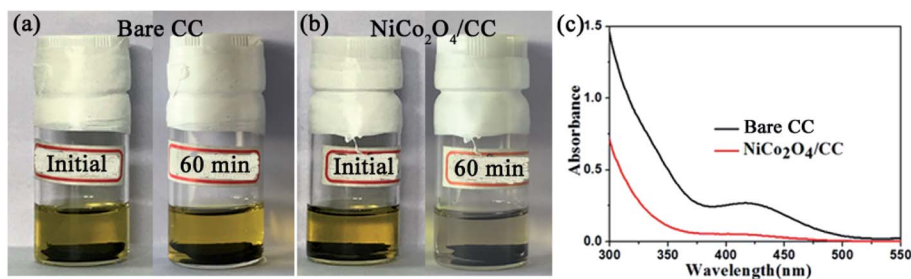


Fig. 6 Polysulfide adsorption properties of CC and NiCo₂O₄/CC. Visual experiments with (a) CC and (b) NiCo₂O₄/CC placed in the solution of Li₂S₄/DOL/DME (2 mmol, DOL : DME volume ratio 1 : 1) at the beginning and after 60 min. (c) UV-vis absorption curves of Li₂S₄/DOL/DME solution after 1 h adsorption by bare CC and NiCo₂O₄/CC.



Further, the superior battery performance including higher specific capacity, enhanced cycling stability and improved rate performance for the S/NiCo₂O₄/CC cathode compared to the S/CC cathode is demonstrated in Fig. 5e. When cycled at 0.1, 0.2, 0.5, 1 and 2C, the discharge capacities of the S/NiCo₂O₄/CC cathode are 1480, to 1200, 1059, 850 and 624 mA h g⁻¹, respectively. In contrast, for the S/CC cathode, the capacities at 1 and 2C are significantly lower, at levels of 300 and 200 mA h g⁻¹, respectively. When the current density returns to 0.1C, a high capacity of 1250 mA h g⁻¹ is achieved, demonstrating the excellent structural stability of the S/NiCo₂O₄/CC cathode at the higher rates of 1C and 2C. The discharge/charge voltage plateaus for the S/NiCo₂O₄/CC cathode at varied rates from 0.1C to 2C are presented (Fig. 5f). The higher the rate, the lower the capacity and the larger the polarization is. When increasing the rate from 0.1C to 1C, the discharge voltage plateau, particularly the second one, becomes lower and shortens. At 2C, the second discharge voltage plateau almost disappears and the capacity keeps decreasing. Similarly, from 0.1 to 2C, the charge voltage plateau increases from around 2.3 to 2.5 V, and shortens. Moreover, the EIS measurements for S/CC and S/NiCo₂O₄/CC after cycling at 100% charge state were carried out and are shown in Fig. S10.† The EIS spectra of S/CC show two depressed semicircles while the EIS spectra of S/NiCo₂O₄/CC are composed of only one depressed semicircle in high frequency regions. As indicated in ref. 50, the semicircle in the middle frequency range is caused by the solid Li₂S (or Li₂S₂) film on the matrix in the cathode; the impedance spectra of the fully charged cathode should display only one semicircle. This indicates that, after recharging, the solid Li₂S does not entirely transform back into the polysulfides for S/CC. However, the

semicircle due to the solid Li₂S was not detected for S/NiCo₂O₄/CC, indicating that the solid Li₂S entirely transformed back into polysulfides. Therefore, the EIS results further confirm that NiCo₂O₄ could act as a catalyst to improve the electrochemical kinetics for polysulfide conversion.

For the Li-S battery cathode, soluble LiPS are formed during the transition between S and Li₂S/Li₂S₂ for both the charge and discharge processes. Hosts that can adsorb LiPS and catalyze/accelerate their electrochemical transformations are therefore needed, because they are capable of improving the performance by retaining LiPS against diffusion towards the anode and enhancing the charge transfer kinetics. In light of these considerations, Fig. 6 depicts the polysulfide adsorption properties of NiCo₂O₄/CC; Fig. 7 shows the boosted kinetics of the S/NiCo₂O₄/CC cathode, and Fig. 8 illustrates the mechanism of the improved battery performance.

Polysulfide adsorption of NiCo₂O₄/CC was revealed with a visual experiment and UV-vis absorption test. CC and NiCo₂O₄/CC was separately placed in the yellow solution of Li₂S₄. After 60 min, the Li₂S₄ solution with NiCo₂O₄/CC faded to colorless while the other with CC stayed unchanged (Fig. 6a and b), demonstrating the superior adsorption properties of the polar NiCo₂O₄ nanosheets. The nonpolar CC has no such adsorption ability. Besides the polarity of NiCo₂O₄, the high surface area is another determining factor for the excellent adsorption of LiPS. The pores ranging from micropores to macropores offer abundant adsorption sites to entrap the soluble LiPS on NiCo₂O₄ nanosheets. The supernatant was then analyzed with UV-vis absorption (Fig. 6c). The characteristic absorption peak of S₄²⁻ at 410 nm remains for CC, but disappears for NiCo₂O₄/CC. The strong polysulfide adsorption of

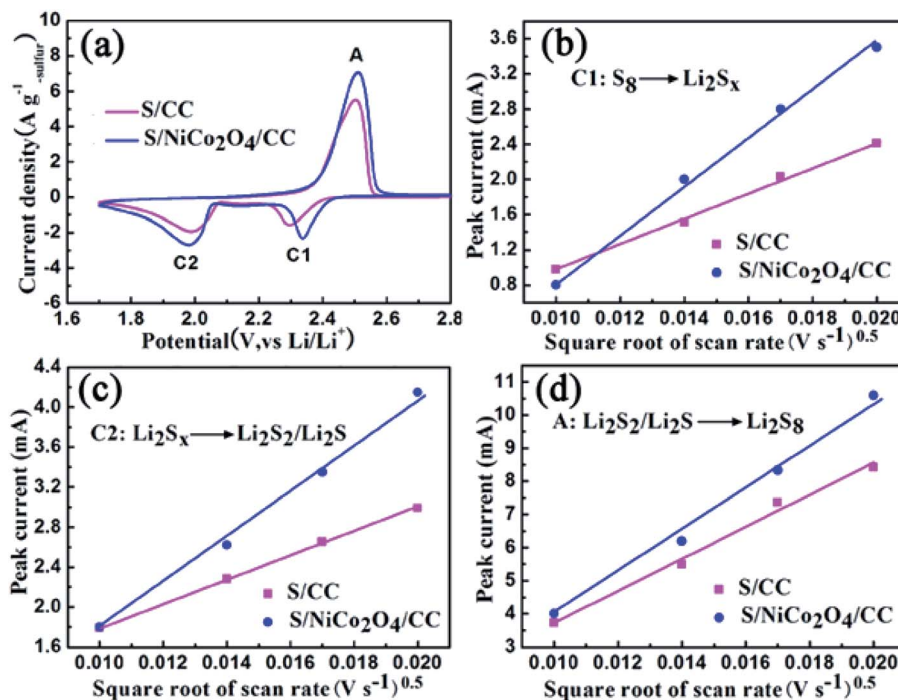


Fig. 7 (a) CV curves for S/CC and S/NiCo₂O₄/CC at 0.4 mV s⁻¹. (b–d) Plots of the peak current against the square root of scan rate for peaks C1, C2 and A.



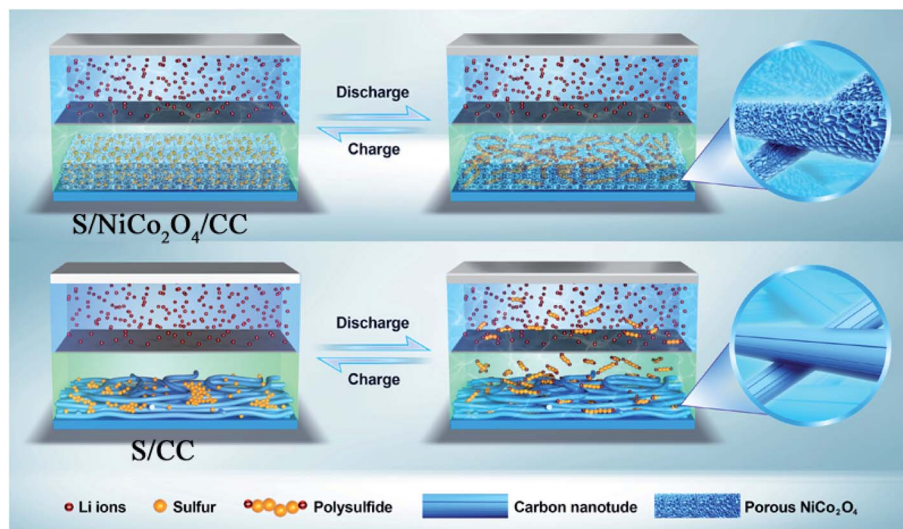


Fig. 8 Schematic illustration of the charge/discharge properties of the Li-S cells with S/NiCo₂O₄/CC and S/CC cathodes.

NiCo₂O₄/CC enables the higher capacity and enhanced cycle stability of S/NiCo₂O₄/CC compared to S/CC.

In a typical CV comparison, the cathodic peaks located at 2.32 (peak C1) and 2.10 V (peak C2) correspond to the gradual reduction of sulfur to LiPS and then to Li₂S₂/Li₂S (Fig. 7a). The sharp peak located at 2.50 V (peak A) in the anodic scan represents the reverse conversion. The peak currents for S/NiCo₂O₄/CC are larger than those for S/CC. Besides, the cathodic peak C1 for S/NiCo₂O₄/CC is at a higher potential than that for S/CC. These characteristics indicate the improved electrochemical kinetics for S/NiCo₂O₄/CC.

To further study the electrochemical kinetics, CV measurements were carried out with scan rates of 0.1, 0.2, 0.3 and 0.4 mV s⁻¹ (Fig. S11†). The maximum currents for the three peaks of C1, C2 and A are plotted against the square root of scan rates separately (Fig. 7b–d). They all possess a linear relationship, with the slope representing the diffusion rate of lithium ions, according to the Randles–Sevcik equation (ESI†).^{51,52} Meanwhile, for each redox peak, the slope for S/NiCo₂O₄/CC is larger than that for S/CC, indicating the accelerated diffusion of LiPS owing to the interconnected NiCo₂O₄ nanosheet arrays on the CC fibers. The adsorption and catalytic properties of NiCo₂O₄ nanosheets enhance the transformation rates of sulfur species, thereby boosting the diffusion of LiPS. NiCo₂O₄ as a catalyst accelerates the charge-transfer process and improves the electrochemical kinetics.

For the S/NiCo₂O₄/CC cathode, the interconnected NiCo₂O₄ nanosheet arrays act as a polar host, absorber and catalyst for the sulfur species, and the CC fibers enhance the electrical conductivity. As a host, the distinctive structure and the high surface area of NiCo₂O₄ cause homogeneous dispersion of elemental sulfur, as well as the discharge products and intermediates. The polar feature of NiCo₂O₄ attracts soluble LiPS, thus retaining them and preventing their diffusion towards the anode side. In contrast, the nonpolar CC fiber has no such effect (Fig. 8). The combined effects of the interconnected NiCo₂O₄ nanosheet arrays therefore endow the S/NiCo₂O₄/CC cathode

with greatly improved electrochemical performances compared to S/CC. As shown by demonstration, a Li-S coin cell with the S/NiCo₂O₄/CC cathode is capable of illuminating LED lights for 6 h (Fig. S12†).

3. Conclusion

In summary, interconnected NiCo₂O₄ nanosheet arrays on carbon cloth fibers are proposed as a polar host, adsorber and catalyst for sulfur species. As a host, the distinctive structure and the high surface area of NiCo₂O₄ cause homogeneous dispersion of elemental sulfur, as well as the discharge products and intermediates. The polar feature of NiCo₂O₄ attracts soluble LiPS, thus retaining them and preventing their diffusion towards the anode side. In contrast, the nonpolar carbon fiber has no such effect. NiCo₂O₄ as a catalyst accelerates the charge-transfer process and improves the electrochemical kinetics. Consequently, the prepared S/NiCo₂O₄/CC electrodes show a high initial capacity of 1480 mA h g⁻¹ at 0.1C and enhanced rate capability (850 and 624 mA h g⁻¹ at 1C and 2C, respectively), as well as stable cyclability (822.3 mA h g⁻¹ after 400 cycles at 0.5C with a low capacity decay rate of 0.060% per cycle).

4. Experimental section

Synthesis of NiCo₂O₄/CC

Co(NO₃)₂·6H₂O (0.5 mmol), Ni(NO₃)₂·6H₂O (1 mmol), and Hexamethylenetetramine (3 mmol) were added into an autoclave with methanol (24 mL) to form a homogeneous light pink solution under stirring. Then as-cleaned carbon cloth was vertically placed into the autoclave, which was heated at 180 °C for 12 h. The as-treated carbon cloth was washed and annealed at 350 °C for 3 h to obtain NiCo₂O₄/CC.

Synthesis of S/NiCo₂O₄/CC

S/NiCo₂O₄/CC was prepared by sulfur impregnation into NiCo₂O₄/CC followed by melt diffusion. The sublimed sulfur



was first dissolved in carbon tetrachloride (CCl_4) and then the as-obtained $\text{NiCo}_2\text{O}_4/\text{CC}$ was immersed into the sulfur-containing solution. After immersing for a certain period of time, the piece was taken out for drying. The sulfur loading amount of $\text{S}/\text{NiCo}_2\text{O}_4/\text{CC}$ was controlled by specifying the immersion time and the concentration of the S/CCl_4 solution. In order to achieve better dispersion of sulfur in $\text{S}/\text{NiCo}_2\text{O}_4/\text{CC}$, the melt diffusion process ($155\text{ }^\circ\text{C}$ for 12 h) was further applied. Following a similar procedure, the control sample S/CC was prepared.

Polysulfide adsorption experiment

Li_2S and S at a molar ratio of 1 : 3 were added into DOL/DME (volume ratio 1 : 1) under stirring at $60\text{ }^\circ\text{C}$ for 12 h to form a homogeneous Li_2S_4 solution (2 mmol L^{-1}). $\text{NiCo}_2\text{O}_4/\text{CC}$ and bare CC were added into the Li_2S_4 solution (3 mL) and allowed to rest for 1 h, respectively. Then the supernatants were analyzed with a UV-Vis Spectrophotometer (UV-Vis, Shimadzu, UV2450).

Electrochemical measurements

The S/CC and $\text{S}/\text{NiCo}_2\text{O}_4/\text{CC}$ pieces were cut into a circular shape with a diameter of 14 mm to use as the cathodes of Li-S batteries. The integrated cathode, the lithium metal anode and a Celgard 2300 separator were assembled into CR2032 coin cells in an argon-filled glove box with both moisture and oxygen contents below 0.1 ppm. The electrolyte was composed of LiTFSI (1 M) and LiNO_3 (0.2 M) in DME : DOL (1 : 1, v/v). Galvanostatic discharge-charge tests of S/CC and $\text{S}/\text{NiCo}_2\text{O}_4/\text{CC}$ were carried out on a Battery Measurement System (Shenzhen Neware Electronic Co., China) with a voltage window of 1.7–2.8 V versus Li^+/Li . The current density ($1.0\text{C} = 1675\text{ mA g}^{-1}$) and the specific capacity were calculated based on the mass of the loaded sulfur. If not specified, the sulfur loading was controlled at $1.1\text{--}1.3\text{ mg cm}^{-2}$. CV curves were obtained at different scan rates from 0.1 to 0.4 mV s^{-1} between 1.7 and 2.8 V (vs. Li/Li^+). Electrochemical impedance spectroscopy (EIS) tests were carried out at 5 mV amplitude with the frequency from 100 kHz to 10 mHz.

Materials characterization

X-ray diffraction (XRD) tests were performed between 10° and 80° at 40 kV and 50 mA on a Bruker D8 ADVANCE X-ray Diffractometer. The microscopic morphology and structure were determined on an SU8820 Hitachi field-emission scanning electron microscope (FESEM). TG analysis was conducted on a DSC3+ Mettler Toledo. The sample was heated in a nitrogen atmosphere at $10\text{ }^\circ\text{C min}^{-1}$ from room temperature to $600\text{ }^\circ\text{C}$. BET analysis was carried out on ASAP2020 Micromeritics Instruments.

Conflicts of interest

The authors declare no conflict of interest.

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