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Anisole hydrodeoxygenation over Ni–Co bimetallic catalyst: a combination of experimental, kinetic and DFT study†

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Catalytic hydrodeoxygenation (HDO) of anisole was performed with a series of Ni and Co containing catalysts with different weight ratios on activated carbon (AC) for cyclohexanol production. The catalytic activities of various catalysts revealed that Ni₅Co₅-AC was the best catalytic system. Structural analysis obtained from XRD, TPR, XPS, and TEM evidently demonstrates that Ni₅Co₅-AC sample consists of a distorted metal alloy spinel structure and optimum particle size, enhancing its catalytic performance. Kinetics were investigated to identify cyclohexanol production rate, activation energy, and reaction pathway. Structural, experimental, kinetics and density functional simulations suggested that high amount of distorted metallic alloy in Ni₅Co₅-AC, presence of water, high adsorption efficiency of anisole, and low adsorption tendency of cyclohexanol on metallic alloy surface were the critical factors for HDO of anisole to cyclohexanol.

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1. Introduction

Selective transformations of lignin derived compounds into aliphatic cyclic alcohols are attractive and fascinating for the future because these bio-based aliphatic cyclic alcohols are building blocks of amine derivatives, carboxylic acids, ketones, resins, polymers, paints, fine chemicals, and rubber.¹ However, the selective hydrodeoxygenation process to produce cyclic alcohol is challenging. It includes the hydrogenation of unsaturated functional groups such as C=C and C=O and rupturing of O–C bond.² Conventional hydrogen and liquid organic hydrogen carriers (LOHCs) can be used for HDO.³ However, noble metal catalysts, high temperature and uncontrolled HDO limit the use of LOHCs for low cost selective production of cyclic alcohol.^{4–6}

Non-noble metal-based catalytic systems with high hydrogenation capability and controlled deoxygenation are critical for developing an efficient process.^{7–10} Literature reports suggest that the combination of Ni with transition metals such as Fe, Mo, Co, and Cu is more active and selective for HDO under ambient reaction conditions.^{3,4} Therefore, numerous bimetallic catalysts Ni–Co, Ni–Fe, Ni–Cu, and Ni–Mo have been used for HDO. It was found that Ni–Mo and Ni–Fe based catalysts are suitable for the production of aromatic compounds.^{5,6} Huynh *et al.*¹¹ compared the activity of Ni–Co/ZSM-5 and Ni–Cu/ZSM-5 for phenol HDO and identified Cu incorporation to Ni deteriorated the catalytic performance owing to variable and large sized Cu species. At the same time, Co increased the activity along with selectivity. Co insertion promoted small and strongly stabilized active metal alloy sites with outstanding H₂ activation and dissociation capability. Similarly, Liu and co-workers identified Co to be highly active for the selective cleavage of C_{aryl}–OCH₃ bond.^{12,13}

Metal compositions significantly affect the textural property, reducibility, and acidity of the Ni–Co bimetallic¹⁴ catalysts. These changes are reflected in their catalytic performances. Ni–Co nanoparticle decorated porous Carbon/ZrO₂ with different ratios of metals were used for phenol conversion.¹⁵ Ni₃Co₁@C/ZrO₂ (15 wt%) showed the highest conversion (96%) and cyclohexanol yield (91%). The Ni–Co interaction and formation of bimetallic alloy were main causes for high activity. Zhou *et al.*¹⁶ investigated the hydrodeoxygenation of guaiacol over Ni–Co bimetallic systems with different ratios of Ni and Co. The equal metal loadings showed better activity than other compositions, taking

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temperature, hydrogenation and catalyst acidity into consideration. In another attempt, same research group investigated the reasons of high activity and reported that proper acidity, good metal dispersion and interaction between metal particles and support were responsible for it.¹⁴ Similarly, Li *et al.*¹⁷ investigated the impact of Co/Ni molar ratio (1:0 to 0:1) by changing incorporation amount of Ni and Co. Cyclohexanol amount in reaction system was increasing with Ni content, became maximum at 1Co–1Ni and then started decreasing. Authors claimed lowest alloy particle size to be a crucial factor for such high activity. Graphite supported bimetallic Ni–Co catalysts demonstrated improvement in reducibility, smaller particle size, and stability than monometallic catalyst.¹⁸ High Co containing catalyst led to higher selectivity towards phenol, while Ni₃Co₁ was the most active for cyclohexanol ($\approx 65\%$) production with $\approx 80\%$ guaiacol conversion. Very recently, Wang and co-workers¹⁹ compared the activity of Ni_xCo_y bimetallic catalysts (Ni_x:Co_y = 1:0, 2:1, 1:1, 1:2, 1:3, 0:1) for cyclohexanol production. Results suggested that highest synergistic effect between Ni and Co, surface area and uniformly dispersed Ni–Co bimetallic sites in NiCo(1/2)@C-600 lead selectivity.

Although mentioned literature shows some promising signs for hydrodeoxygenation progress for cyclohexanol production, however, in-depth understanding of kinetic and mechanisms of the process and appropriate explanation of real time behaviour of Ni–Co bimetallic catalyst is still lacking. It is very crucial for a step forward to tune an efficient and engineered Ni–Co bimetallic catalytic systems with high hydrogenation capability and controlled deoxygenation activity for industry relevant cyclohexanols production. To contribute in this direction, herein, we developed Ni–Co nanoparticles impregnated activated carbon systems for selective HDO of anisole. The main objectives of this study are to understand effects of metal composition on active sites, and their role according to kinetic aspects with mechanistic pathways. The optimized catalyst showed unique behavior in both hydrogenation activity and deoxygenation selectivity. The kinetic and theoretical models were investigated to understand Ni–Co behaviour in real time reaction system to validate the mechanistic pathways for anisole hydrodeoxygenation.

2. Material synthesis and method

2.1. Materials

Activated carbon, ethyl acetate and other solvents were purchased from Fischer Scientific ($\geq 99\%$ purity). Anisole, and metal precursors, all are $\geq 98\%$ purity, were purchased from Sigma-Aldrich and were used as received without any further treatment. The used double distilled water was synthesized in lab.

2.2. Catalyst synthesis

Activated carbon supported 10 wt% Ni, and Co catalysts were prepared by simultaneous co-impregnation method according to our previous study.¹ The different loadings of Ni and Co were done by dissolving the calculated amount of Ni(NO₃)₂·6H₂O

and Co(NO₃)₂·6H₂O in water (slightly in excess of the pore volume of the AC). The resultant slurry was stirred at room temperature with 200 rpm for 1 h then it was stirred at 80 °C for 2 h. The resultant catalysts were dried overnight at 105 °C, subsequently calcined at 325 °C in the presence of continuous nitrogen flow (100 mL min^{−1}) with the 10 °C min^{−1} ramping rate for 2 h after reaching to the desired temperature. The calcined catalysts were reduced in continuous hydrogen flow (100 mL min^{−1}) for 4 h at 600 °C. The reduced and powdered catalysts were used for characterization and catalytic activity.

2.3. Catalyst characterization

Prepared catalysts were characterized with various techniques such as powder X-ray diffraction (P-XRD), field emission scanning electron microscopy (FE-SEM) with energy-dispersive X-ray spectroscopy (EDS), transmission electron microscopy (TEM), N₂-physisorption, H₂-temperature programmed reduction (H₂-TPR), NH₃-temperature programmed desorption (NH₃-TPD), Fourier-transform infrared spectroscopy (FT-IR), thermo-gravimetric analysis (TGA) and X-ray photoelectron spectroscopy (XPS).

3Flex Physisorption, Micromeritics was used for N₂-physisorption at liquid nitrogen temperature (−196 °C) to calculate the surface area, pore-volume, and pore size of the catalysts. Prior to the analysis, the surface of prepared catalysts was cleaned at 350 °C for 4 h under 1×10^{-5} torr vacuum. Brunauer–Emmett–Teller (BET) equation was used to determine the specific surface area (S_{BET}) from the obtained adsorption data ($P/P_0 = 0.05$ – 0.25).²⁰ The average pore size was reported as a maxima of pore size distributions (PSDs) and these PSDs were evaluated from Barrett–Joyner–Halenda (BJH) algorithm using nitrogen adsorption branch data.²¹ The pore volumes of as prepared catalysts were measured as adsorbed volume of liquid nitrogen at $P/P_0 \approx 1$.

H₂-TPR, chemisorption and NH₃-TPD were employed on Micromeritics, Auto Chem II-HP 2920 in a U shaped quartz tube reactor equipped with a thermal conductivity detector (TCD). Prior to H₂-TPR, chemisorption and NH₃-TPD, the samples were heated with 10 °C min^{−1} in Ar and He, respectively to 300 °C for pre-treatment and temperature was maintained for 1 h. The pre-treated samples were cooled to 50 °C and analysis was done in the temperature range of 50–900 °C with a heating rate of 10 °C min^{−1}. For NH₃-TPD, the 50 mL min^{−1} flow of 10% NH₃–He for 30 min was used for adsorption of NH₃ on pretreated samples. The NH₃ adsorbed samples were purged with pure He (20 mL min^{−1}) for 30 min to remove the extra NH₃–He. Subsequently, the temperature of prepared sample was raised to 900 °C and desorbed NH₃ molecules were monitored on TCD. For H₂-TPR, the analysis of pre-treated samples (non-reduced) was performed in 10% H₂–Ar for the temperature range of 50–900 °C. IPA-liquid nitrogen slurry was used for condensation of released water vapors in the trap area, and quantity of adsorbed H₂ was recorded in the form of TPR signals on TCD.

XRD patterns of powder samples were recorded in the range of $2\theta = 2$ – 80° on a Bruker D8 advance X-ray diffractometer fitted with a Lynx eye high-speed strip detector. The Cu K α ($\lambda = 1.5418$ Å) used as a radiation source, while the scan rate was 0.02 step

per sec for all samples. Shimadzu DTG-60 was used for TGA of catalysts. For TGA, the samples were heated in alumina cell with $10\text{ }^{\circ}\text{C min}^{-1}$ under inert atmosphere (100 mL min^{-1} flow of N_2) from room temperature to $900\text{ }^{\circ}\text{C}$. FT-IR of samples (prepared with KBr) were recorded in the range of $400\text{--}4000\text{ cm}^{-1}$ using FTIR-Perkin Elmer-Spectrum II instrument.

TEM/HRTEM was performed on JEOL JEM 2100 microscope. The images were taken by mounting an ethanol-dispersed sample on a lacey carbon formvar Cu grid at an operating voltage of 200 kV . Elemental mapping was also recorded using same spectrophotometer and EDS was used for the elemental composition of samples. SEM images were taken with FEI Quanta 200 F, having tungsten filament as an X-ray source doped in lanthanum hexaboride (LaB_6), fitted with an ET (Everhart-Thornley) detector, using secondary electrons and an acceleration tension of 10 or 30 kV in high vacuum.

XPS for the surface study of Ni-Co catalyst was performed, using ESCA+, (omicron nanotechnology, Oxford Instrument Germany) equipped with monochromator aluminum source (Al $K\alpha$ radiation $h\nu = 1486.7\text{ eV}$). The instrument was operated at 15 kV and 20 mA . To overcome the charging problem, a charge neutralizer of 2 keV was applied, and binding energy of C $1s$ core (284.6 eV) was taken as reference. High-resolution XPS (HR-XPS) was collected by passing energy of 69.0 eV with a step size of 0.125 eV . The HR-XPS profiles of Ni and Co were fitted using a Gaussian function, and peak positions were normalized corresponding to the C $1s$ (284.6 eV).

2.4. Catalytic activity

The catalytic activity of various catalysts was performed in a 30 mL customized stainless steel tubular reactor equipped with a pressure gauge. The 100 mg reduced catalyst, 0.5 mmol anisole and 5 mL water were loaded in the reactor. The reactor was purged thrice with 1 MPa hydrogen for generating hydrogen environment and then pressurized to desired hydrogen pressure. The pressurized reactor was fitted to the pre-heated silica oil bath with magnetic stirring at 1000 rpm for the reaction time. The reactor was depressurized after cooling in water bath and organic fraction was collected through liquid-liquid separation. Ethyl acetate (15 mL) was used for extraction of organic phase and ethyl acetate was evaporated from organic mixture in rotatory evaporator. The obtained organic fraction was diluted with 3 mL ethyl acetate and collected in 5 mL vial.

The qualitative analysis of the organic fraction was performed in Agilent GC (7890B) equipped with a mid polar capillary column (DB-35MS, 35% phenyl/ 65% dimethylpolysiloxane, $30\text{ m} \times 0.32\text{ mm} \times 0.25\text{ }\mu\text{m}$) connected to a mass spectrometer (5977A MSD). The quantitative analysis of the product was performed on Shimadzu GC-FID (GC-2014) equipped with RXi-MS (100% dimethylpolysiloxane, $30\text{ m} \times 0.25\text{ mm} \times 0.25\text{ }\mu\text{m}$) capillary column. *n*-Dodecane was used as internal standard for quantification. Conversion and selectivity is reported as the average of two reactions, performed at the same experimental conditions to minimize experimental uncertainty. The conversion and selectivity are reported with $\pm 2\%$ standard deviation and calculated as follows

$$\text{Conversion (\%)} = \left| \frac{n_{\text{reactant}}^0 - n_{\text{reactant}}^f}{n_{\text{reactant}}^0} \right| \times 100 \quad (1)$$

$$\text{Selectivity (\%)} = \left| \frac{n_{\text{sh}}}{n_{\text{reactant converted}}} \right| \times 100 \quad (2)$$

where, n_{reactant}^0 is the initial concentration of reactant and n_{reactant}^f is the concentration of reactant after reaction in mole per litre. The n_{sh} denotes the mol per litre concentration of selective hydrodeoxygenated product.

2.5. Computational modelling

To model the catalytic reactions, spin-polarized periodic density functional theory (DFT) calculations were carried out using Quantum Espresso (QE) package.^{22,23} The PBE (Perdew-Burke-Ernzerhof) exchange-correlation density function²⁴ was used in the calculations, while the core electrons were represented by the Vanderbilt ultrasoft pseudo potential.²⁵ The Kohn-Sham orbitals were expanded with plane-wave basis sets with a kinetic energy cutoff of 300 eV , and the k -point sampling used for the integration in reciprocal space was limited to Gamma-point only. The experimental characterization revealed that the catalyst has a chemical composition of $\text{Ni}_5\text{Co}_5\text{-AC}$ and the formation of (111) surface. Thus, to account for the metal composition effects on the reaction path, a 4-layered 3×3 , Ni-Co (111) bimetallic surface ($\text{Ni}_{72}\text{Co}_{72}$) was created as shown in Fig. 5. In our model, the activated carbon support was not considered since it does not take part in the catalytic reactions. A vacuum of $15\text{ \AA}$ was added along the Z -axis to minimize the interaction of the surface with its own replica. The structures were optimized using the BFGS algorithm²⁶ as implemented in the QE package, where the top two out of four atomic layers of the 3×3 Ni-Co (111) surface were relaxed, whereas the atoms in the lower two layers were kept frozen to the bulk lattice position.

The adsorption energies (ΔE) of all the species, which could form during the HDO process, were calculated according to the following equation,

$$\Delta E = E_{\text{system}} - (E_{\text{adsorbent}} + E_{\text{adsorbate}}) \quad (3)$$

where, E_{system} is the total energy of adsorbed species on the Ni-Co surface, while $E_{\text{adsorbent}}$ is the energy of adsorbent molecules (e.g. anisole, cyclohexanol, and methoxycyclohexanol (MCH)), and $E_{\text{adsorbate}}$ is the energy of bare NiCo surface.

3. Results and discussion

3.1. Catalyst characterization

Physical characteristics of catalysts are given in Table S1.† Table states that surface area and pore volume of AC are $938\text{ m}^2\text{ g}^{-1}$ and $0.67\text{ cm}^3\text{ g}^{-1}$, respectively. In all prepared catalysts, there is decrease in surface area and pore volume due to deposition of metal particles.²⁷ There is no drastic difference in surface areas of all prepared catalysts. XRD patterns of activated carbon supported Ni and Co catalysts are shown in Fig. 1a. A broad peak in the region $21\text{--}29^{\circ}$ centred around at 26° is attributed to carbon. The $\text{Ni}_{10}\text{Co}_0\text{-AC}$ and $\text{Ni}_0\text{Co}_{10}\text{-AC}$ show the typical



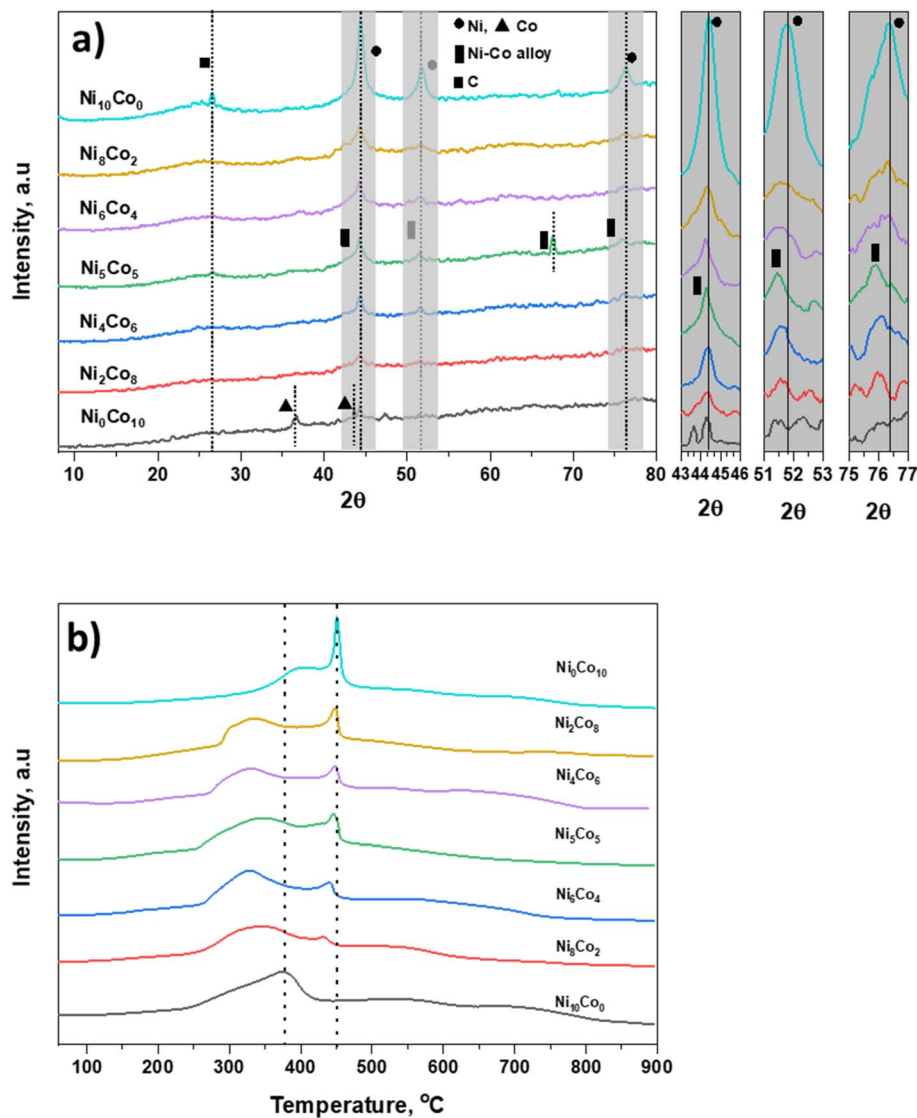


Fig. 1 (a) P-XRD pattern; (b) H_2 -TPR profile of activated carbon supported NiCo-AC catalysts.

diffraction patterns corresponding to Ni and Co metal.^{28,29} The most intense diffraction peaks of Ni and Co are located at 2θ value of 44.4° (111), and 44.2° (111), respectively.¹⁴ These characteristic peaks of Ni and Co do not precisely match with the Ni-Co bimetallic catalysts. The peak of Ni-Co bimetallic catalyst shifted corresponding to Ni-Co alloy (111), can be seen in inset figure between 43 and 46° .^{30,31} The other signature peaks of nickel phase (51.8° , and 76.3°) also shift towards lower 2θ values (51.5° and 75.8°) with reduced intensities associated with the Ni-Co alloyed phase (200) and (220),^{30,32} respectively. All these peaks reposition from their original positions, suggesting the formation of Ni-Co alloy due to the metal-metal interaction. Besides this, an additional peak can be seen at 67.5° corresponding to Ni-Co alloy phase in Ni_5Co_5 -AC.³⁰ Therefore, Ni-Co metal alloy formation in the Ni-Co bimetallic catalyst depends on the Ni/Co ratio. These results are well in accordance with the previous literature reports, which suggested the existence of the characteristic diffraction peak in the region (43 – 47°) and showed the peak shift in between 51 – 53°

and 75 – 77° for the supported Ni-Co bimetallic catalysts due to the formation of Ni-Co alloy phase.^{30,33–35}

H_2 -TPR of all catalysts was employed for the identification of the reducibility of catalyst and the specific interaction between metal species and support (Fig. 1b). TPR spectra of $\text{Ni}_{10}\text{Co}_0$ -AC consists of one broad reduction peak in the range of 250°C to 425°C , centred at 375°C . TPR spectra of $\text{Ni}_0\text{Co}_{10}$ -AC consists of one large peak at 450°C with a shoulder peak at 296°C , it reflects two-step reduction of Co_3O_4 to CoO and subsequent CoO reduction to Co .³⁶ High reduction temperature of $\text{Ni}_0\text{Co}_{10}$ -AC is suggesting a strong interaction between Co species and carbon support. There are two peaks in the mixed oxides. In mixed oxides, a shift in parent peaks to lower temperatures in comparison of $\text{Ni}_{10}\text{Co}_0$ -AC and $\text{Ni}_0\text{Co}_{10}$ -AC indicates that the interaction of Ni and Co atoms in bimetallic catalysts would contribute to the improvement of reducibility of catalysts.^{37,38} This type of interaction has been well reported in literature for Ni/Co catalysts which lead alloy formation on the surface of the support.^{34,39} This peak is at lowest temperature in Ni_5Co_5 -AC



and suggests that this particular sample is highly reducible due to high interaction between Ni and Co. With this point and according to XRD, and TEM results the formation of mixed spinel phase and high dispersion of metals in Ni₅Co₅-AC are confirmed.

Surface texture and microstructure of the synthesized materials were checked using HR-TEM. High magnified TEM images (Fig. 2b, e and h) indicate Ni, Co and mixed metal nanoparticles are embedded on AC surface. The dark spots in TEM images (Fig. 2a, d and g) of catalysts and TEM-EDS mapping (Fig. S1†) show homogeneous distribution of metal nanoparticles. Ni₆Co₄-AC, Ni₅Co₅-AC and Ni₄Co₆-AC demonstrate common lattice distances 0.225 nm and 0.21 nm corresponding to (111) plane of Ni-Co alloy and cobalt metal.^{40,41}

Ni₆Co₄-AC shows two types of inter planar distances for (111) and (220) planes of Ni validating high Ni amount.⁴² Similarly, Ni₄Co₆-AC surface texture (Fig. 2e) shows cobalt, alloy and mixed oxides (220) phases.⁴³ SAED pattern of Ni₅Co₅-AC represents somehow crystallinity possibly due to high content of Ni-Co alloy. Measured metal particles range and average size (Fig. 2c, f and i) of the Ni₆Co₄-AC, Ni₅Co₅-AC and Ni₄Co₆-AC catalysts suggest that cobalt promotes the size reduction and dispersion.⁴⁴ Average size was minimum (3.2 ± 0.2 nm) for Ni₅Co₅-AC and is in consistent with literature.^{17,44,45}

Chemical state and surface element composition of Ni₆Co₄-AC, Ni₅Co₅-AC and Ni₄Co₆-AC catalysts were analyzed by XPS to further identify the interaction between Ni-Co. The peak intensity of Ni and Co in survey scan indicates Ni and Co

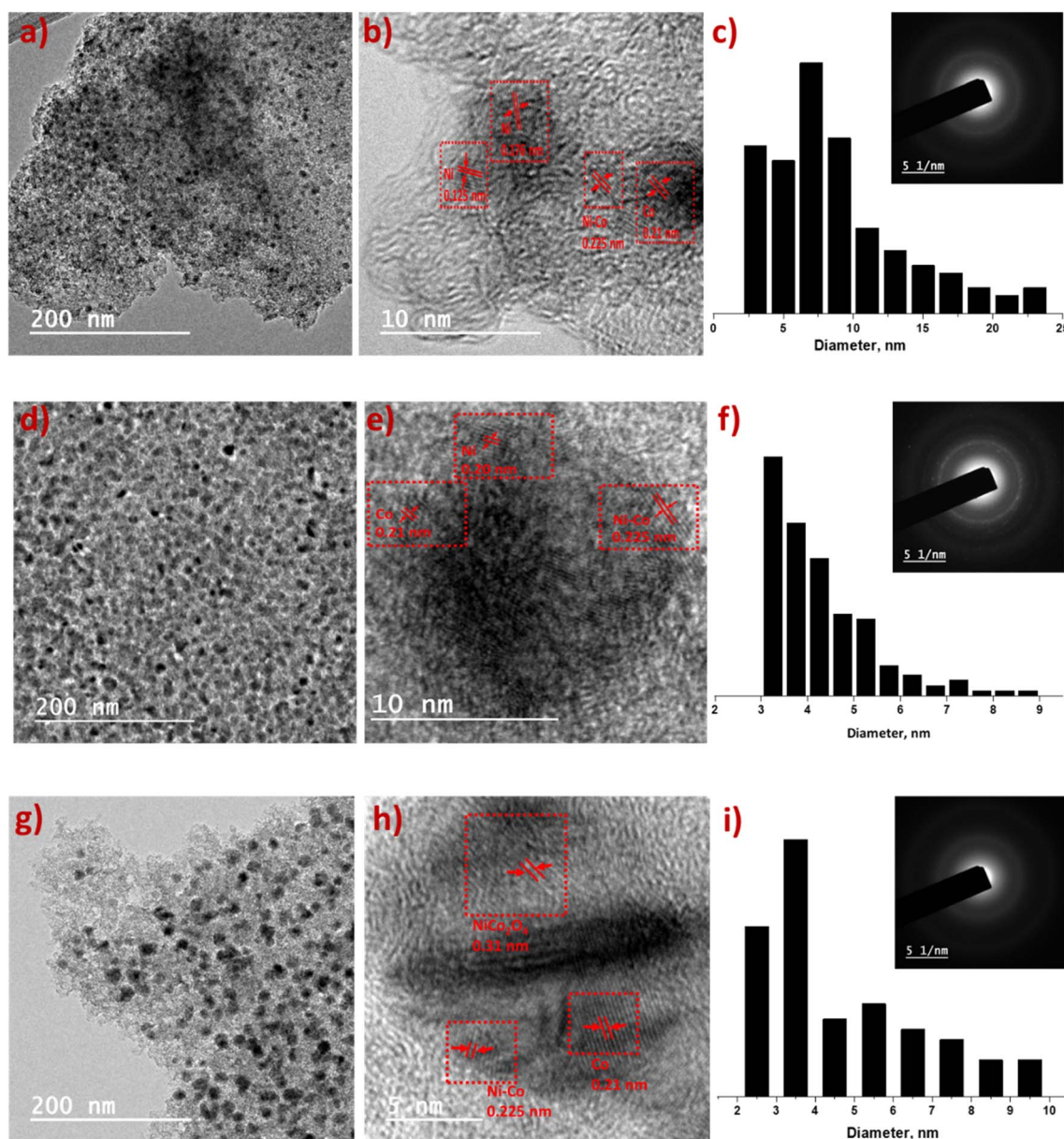


Fig. 2 HR-TEM images of Ni₆Co₄-AC (a–c), Ni₅Co₅-AC (d–f) and Ni₄Co₆-AC (g–i). (b, e and h) interplanar *d*-spacing in high magnified TEM; (c, f and i) particle size range of metals and SAED pattern.



amounts in catalysts (Fig. S2†). In HR-XPS graphs of Ni 2p, the binding energies of Ni 2p_{3/2} and Ni 2p_{1/2} are positively shifted (Fig. 3a) from original peaks (852.4 and 870.1 eV).¹⁴ Likewise in the Co 2p spectra (Fig. 3b), the peak at 777.8 (Co 2p_{3/2}) is negatively shifted by 0.4 eV of Co (778.1 eV), however it arises at its original position in Ni₄Co₆-AC that might be due to the high amount of Co.^{14,46} These peak shifts reveal the electronic interactions between Ni and Co and formation of Ni-Co alloy.⁴⁶ Ni 2p_{3/2} and Co 2p_{1/2} of Ni₅Co₅-AC are well-fitted with two doublets of Ni⁰ and Co⁰ suggesting that easily reducible nature of Ni₅Co₅-AC and presence of high amount of Ni-Co alloy than Ni₆Co₄-AC and Ni₄Co₆-AC. These results are well aligned with XRD, TPR and TEM results.

3.2. Prodigious effects of reaction parameters

Anisole conversion of 30% with MCH (76%) and cyclohexanol (24%) as the primary products was observed over Ni₁₀Co₁₀-AC

catalyst. In contrast, the anisole conversion was 68% and selectivity of cyclohexanol and MCH was 80% and 20%, respectively, over Ni₁₀Co₀-AC. The high conversion of anisole over Ni₁₀Co₀-AC demonstrated the high hydrogenation ability of metallic Ni than Co. Anisole conversion and product distribution over monometallic catalysts (Ni₁₀Co₀-AC and Ni₀Co₁₀-AC), and bimetallic catalysts (NiCo-AC) with various Ni/Co ratios are compared to reveal the bimetallic effect in Ni-Co catalyst (Fig. 4a). As shown in Fig. 4a, the presence of Ni with Co in a single catalyst remarkably affects both the activity and selectivity of the catalyst. With the introduction of nickel content, the NiCo₂O₄ phase is formed due to insertion of Ni²⁺ ions into octahedral sites of Co₃O₄ having spinel structure, and NiCo₂O₄ changed into Ni-Co alloy during reduction. Ni/Co ratio strongly influences the formation of these active surface species and their structure as well as the activity of Ni_xCo_y-AC catalysts. Structural distortion in Ni-Co alloy increases with increase in

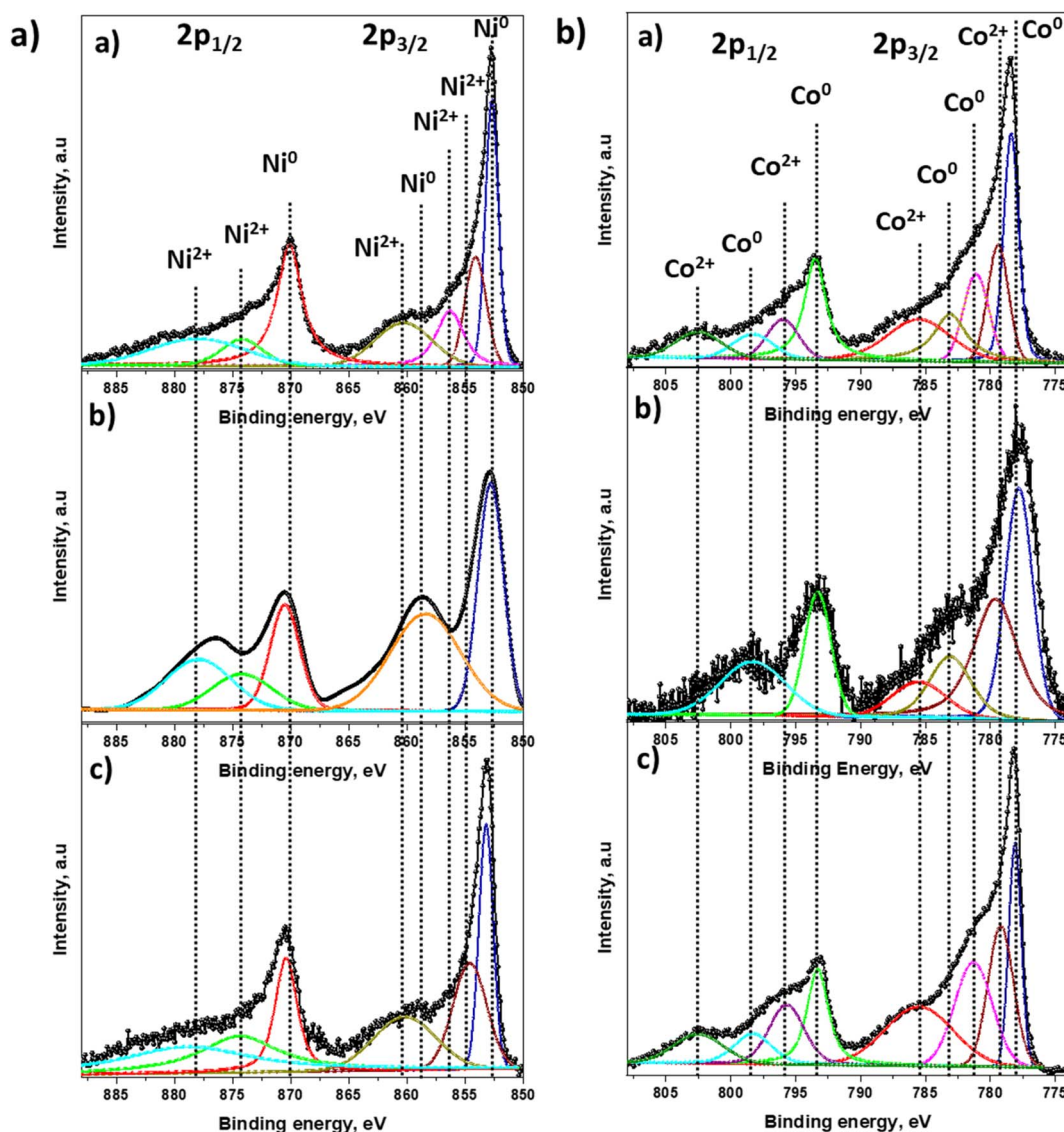


Fig. 3 HR-XPS spectra of (a) Ni 2p; (b) Co 2p. Insight (a)–(c) represents Ni₆Co₄-AC, Ni₅Co₅-AC and Ni₄Co₆-AC.

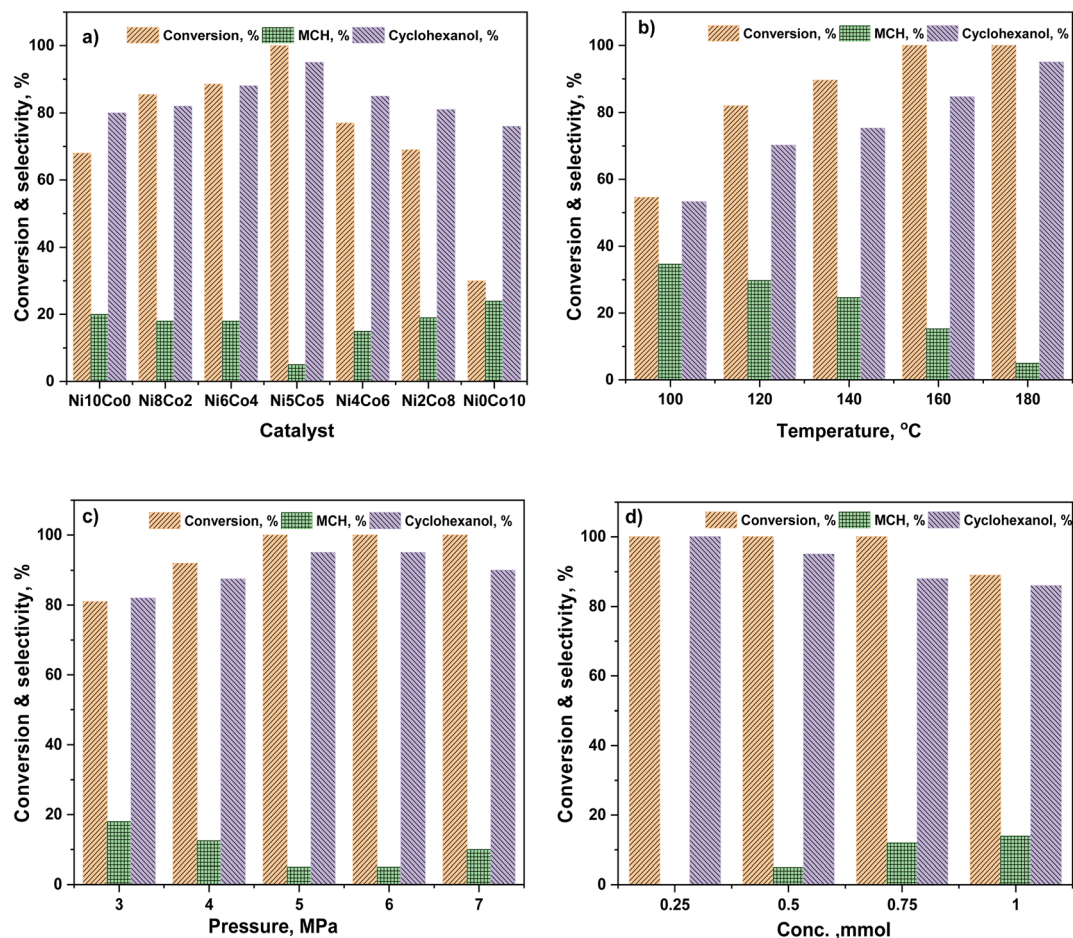


Fig. 4 Anisole hydrodeoxygenation; (a) effect of catalyst composition, (b) effect of temperature, (c) effect of pressure, (d) effect of feed concentration. Reaction conditions: anisole – 0.5 mmol, catalyst – $\text{Ni}_5\text{Co}_5\text{-AC}$ (100 mg), T – 180 °C, P – 5 MPa, water – 5 mL and time – 4 h.

nickel content due to maximum insertion of Ni into octahedral sites.^{47–49} Ni-Co bimetallic sites in $\text{Ni}_x\text{Co}_y\text{-AC}$ works synergistically with the monometallic active sites, and increase the hydrogenation ability.⁴⁷ The high selectivity towards cyclohexanol highlights the critical role of neighbouring Ni-Co bimetallic active sites for the adsorption of anisole and subsequent cleavage of $\text{C}_6\text{H}_{11}\text{O-CH}_3$ bond. Anisole conversion increased to 69% over $\text{Ni}_2\text{Co}_8\text{-AC}$ with 81% cyclohexanol selectivity. High amount of Ni-Co alloy with the highest distortion and small size led complete conversion of anisole over the $\text{Ni}_5\text{Co}_5\text{-AC}$ catalyst, and selectivity towards cyclohexanol (95%) (Fig. 4a). XRD, TEM, $\text{H}_2\text{-TPR}$, and XPS results attest to this fact. A significant drop in anisole conversion (11.5%) and cyclohexanol selectivity (7%) was observed over $\text{Ni}_6\text{Co}_4\text{-AC}$ in comparison to $\text{Ni}_5\text{Co}_5\text{-AC}$ might be owing to low amount of bimetallic sites and large size of metals in $\text{Ni}_6\text{Co}_4\text{-AC}$.

Solvent has a significant effect on anisole conversion and product distribution as observed in various studies.^{50–52} The activity order of $\text{Ni}_5\text{Co}_5\text{-AC}$ for anisole conversion in various solvents is water > isopropanol > isobutanol > ethanol > methanol (Table S2†). The maximum conversion was 13%, and the major product was MCH (72%) when isopropanol was used as

a solvent (entry 3, Table S2†). Results suggest that HDO was a slow reaction in alcohol, and a long reaction time is required for complete conversion (entry 4, Table S2†). However, a drastic change in conversion and product distribution was observed in water where conversion reached to >99%, and cyclohexanol was obtained as a primary product (95%). Low solubility of anisole in water promotes the ability of the reagent to adsorb on active catalyst sites due to less interaction between anisole and water, additionally, high polarity of water causes the hydrogenolysis of ArO-CH_3 bond.^{53,54} This explanation is well confirmed by Yang and co-workers,⁶² who state that hydrogenation of the anisole is expected to take place through the π -complex formation, following to adsorption of anisole on active metal surfaces.

The reaction temperature played a significant role in the HDO of anisole.⁵⁵ The effect of temperature (100–180 °C) for anisole HDO with product distribution over $\text{Ni}_5\text{Co}_5\text{-AC}$ is shown in Fig. 4b. Anisole conversion is only 31%, however, cyclohexanol selectivity was 70.5% at 80 °C. The increase in temperature gradually increased anisole conversion and cyclohexanol selectivity. Conversion reached $\geq 99.9\%$ at 180 °C, and selectivity for cyclohexanol was maximum (95%). On the basis of results, it can be assumed that at high temperature, catalyst



exhibited more demethylation behavior due to scission of $\text{C}_6\text{H}_{11}\text{O}-\text{CH}_3$ bond and cyclohexanol appeared as major product. It might be attributed to the lower bond dissociation energy of $\text{HO}-\text{CH}_3$ (377 kJ mol^{-1} at STP) than $\text{H}-\text{OCH}_3$ ($436.8 \text{ kJ mol}^{-1}$ at STP) and $\text{H}-\text{OH}$ ($498.7 \text{ kJ mol}^{-1}$ at STP).^{56,57} Hydrogen pressure had an enormous effect on both anisole conversion and product distribution (Fig. 4c). Anisole conversion was 81% with 82% cyclohexanol selectivity at 3 MPa. Anisole conversion and cyclohexanol gradually increased with pressure and reached maximum ($>99.9\%$ and 95%) at 5 MPa and no significant change in cyclohexanol selectivity was observed above this pressure. The enhancement in anisole conversion and cyclohexanol selectivity at hydrogen pressure ($\leq 5 \text{ MPa}$) can be clarified from two view points. First, hydrogen was a reactant in the experiment and increased pressure favours the hydrogenation reaction and scission of $\text{O}-\text{CH}_3$ bond. Second, the solubility of hydrogen in water was high at high pressure, which would require improvement of hydrogen accessed on the catalyst active sites during the hydrogenation reaction.^{58,59} The impact of feed conc. can be seen in Fig. 4d. At the lower conc. (0.25 mmol), there was selective conversion towards cyclohexanol. As the conc. of feed was increasing, the selectivity of cyclohexanol was decreasing. It was due to that the reacting molecules were increasing while the active sites were constant. Reusability of catalyst is an important feature of heterogeneous catalysis for economic chemical transformation. In reusability, catalyst was separated by centrifugation after each catalytic run, washed with isopropanol to remove organic adsorbent from the catalyst surface and dried at 105°C . Table S3† shows catalyst reusability results. The catalyst was losing activity in consecutive runs, it was owing to loss and agglomerations of active particles and metal oxide formation^{60,61} (Fig. S3†).

3.4. Mechanistic and kinetic study over $\text{Ni}_5\text{Co}_5\text{-AC}$

Product distribution demonstrates only two products. MCH resulted from the aromatic ring hydrogenation of anisole, which might take place following the π -complex adsorption of the anisole on active sites of the catalyst.⁶² The possible reason for the appearance of cyclohexanol was weak $\text{HO}-\text{CH}_3$ linkage.^{56,57} However, identification of dominant pathway of anisole transformation under hydrotreating conditions is complex.^{56,63} To understand this, time on stream data was recorded (Table S4†). Up to the 10 min reaction time, it was very hard to identify the dominant reaction path. However, MCH was the only product at the very initial stage (5 min). As the reaction time increased, cyclohexanol started appearing due to the demethylation of MCH. It suggests that anisole hydrodeoxygenation follows hydrogenation and subsequent demethylation pathway. DFT calculations were carried out to further validate the reaction mechanism and high selectivity of cyclohexanol. The experimental characterization revealed that the catalyst has a chemical composition of $\text{Ni}_5\text{Co}_5\text{-AC}$ and the formation of (111) surface. Thus, to account for the metal composition effects on the reaction paths, a 4-layered 3×3 , Ni-Co (111) bimetallic surface ($\text{Ni}_{72}\text{Co}_{72}$) was created (Fig. 5), considering catalyst support (activated carbon) as neutral and is not taking part in the HDO of anisole. Similar Ni-Co clusters have been used to study Congo red adsorption.⁶⁴ Heterogeneous catalysis follows catch release mechanism. Thus, the activity of Ni-Co cluster and product distribution depend on the strength of the adsorption of reacting and product molecules. The binding energies are listed in Table 1 for different reaction products including reactant and their binding configurations are shown in Fig. 5. The strong binding energy of anisole (-1.41

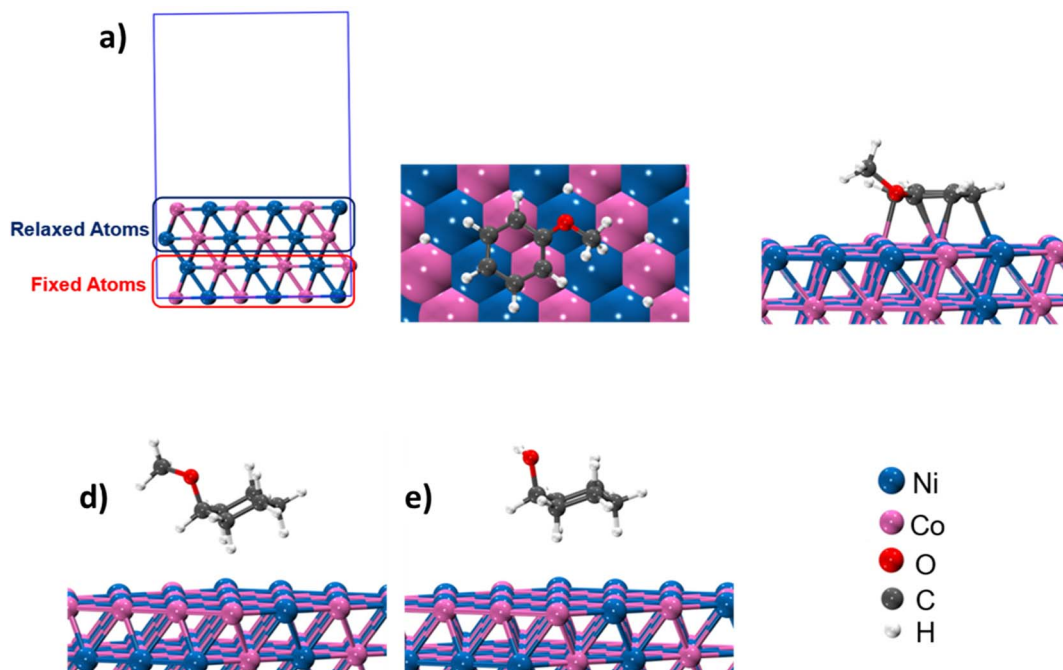
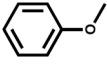
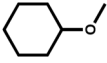
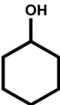
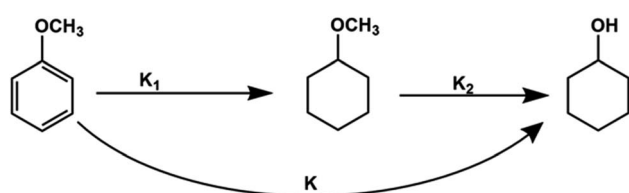


Fig. 5 (a) and (b) Structure and different orientation of $\text{Ni}_5\text{-Co}_5$; (c)–(e) binding configuration of anisole, MCH and cyclohexanol.

Table 1 Binding energy of reactant and product molecules in reaction on the 3×3 surface of Ni–Co (111)

Molecule	Adsorption energy, eV per molecule
	−1.14
	−0.43
	−0.41

eV) is mainly due to the interaction of surface metal atom with the π cloud of the aromatic benzene ring. Since there are no π electrons in the cyclohexane ring, the interaction of the cyclohexanol and MCH molecules to the surface is weak. The binding energy calculations suggested that direct anisole hydrogenation to MCH is kinetically and thermodynamically favourable. It should be noted that MCH can either desorb or undergo further hydrogenolysis to cyclohexanol. MCH appearance in the product profile at very initial stage confirmed its spontaneous desorption on the Ni–Co (111) surface. In contrast, relative low adsorption energy of cyclohexanol assisted the fact of cyclohexanol as the primary product under optimum reaction conditions. The isotope labelling experiment results also suggest the same product distribution although there was deuterated cyclohexanol and MCH (Fig. S4†).

**Fig. 6** Reaction mechanism and pathways of anisole hydrogenation over $\text{Ni}_5\text{Co}_5\text{-AC}$.

For kinetics, reaction route can be simplified (Fig. 6) as an efficient direct reaction from anisole to cyclohexanol due to low selectivity of MCH than cyclohexanol (*i.e.*, $K_1 \ll K_2 = K$). Moreover, there were approx 134 molecules of H_2 for one molecule of anisole and assumed hydrogen in excess and was not considered during calculations.

According to the simplified route, the rate of anisole decomposition can be given as below:

$$-\frac{dC_{\text{anisole}}}{dt} = k \times C_{\text{anisole}}^n \quad (4)$$

where, k and n represent the reaction rate constant and reaction order ($n \neq 1$), respectively.

Integrate the eqn (4) to have

$$C_{\text{anisole}} = [C_{\text{anisole},0}^{1-n} - (1-n) \times k \times t]^{\frac{1}{1-n}} \quad (5)$$

where, $C_{\text{anisole},0}$ represents the original concentration of anisole.

Thus, the real-time concentration of produced cyclohexanol, $C_{\text{cyclohexanol}}$ can be described as follows, *i.e.*,

$$C_{\text{cyclohexanol}} = C_{\text{anisole},0} - [C_{\text{anisole},0}^{1-n} - (1-n) \times k \times t]^{\frac{1}{1-n}} \quad (6)$$

where, $C_{\text{cyclohexanol}}$ represent the real-time concentration of produced cyclohexanol.

According to the Arrhenius equation, the reaction rate constant can be described as:

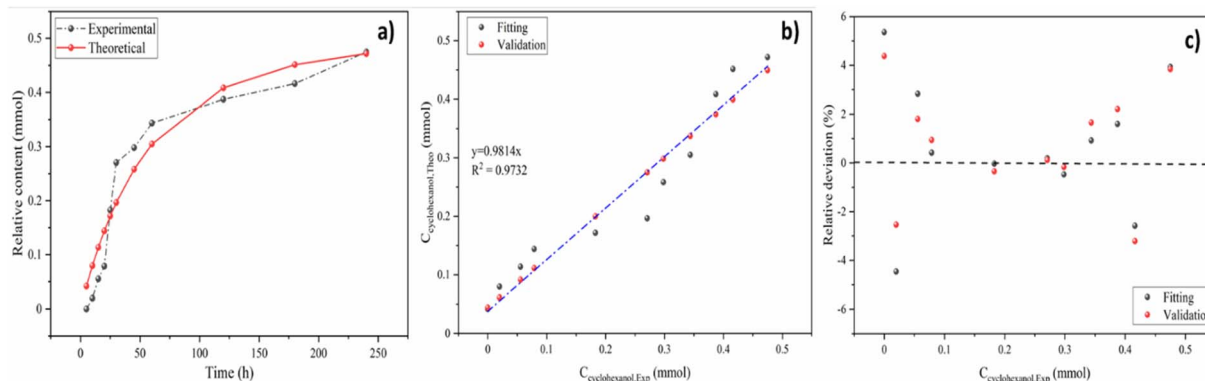
$$\ln k = \ln A - \frac{E_a}{RT} \quad (7)$$

where, E_a and A represents the activation energy and frequency factor, respectively.

According to eqn (6) and (7), the real-time concentration of produced cyclohexanol can be written as

$$C_{\text{cyclohexanol}} = C_{\text{anisole},0} - \left[C_{\text{anisole},0}^{1-n} - A(1-n)e^{\left[-\frac{E_a}{RT}\right]} t \right]^{\frac{1}{1-n}} \quad (8)$$

The real-time concentration measurements of anisole at various temperatures and times were used to compute the

**Fig. 7** The correlation (a), validation (b), and relative deviation (c) between the predicted and experimental data.

parameters of reaction kinetics, *i.e.* E_a , A and n shown in eqn (8). Calculated activation energy and frequency factor through established model were 33.49 kJ mol⁻¹ and 154.38 min⁻¹, respectively. The value of activation energy of cyclohexanol production for Ni₅Co₅-AC is much lower than monometallic Ni-based catalysts (48 and 56 kJ mol⁻¹) and Ni₁₅Co₅@C/ZrO₂ (44.31 kJ mol⁻¹).^{15,65} The reaction order (1.26) is nearly 1, indicating anisole hydrodeoxygenation is a pseudo first-order process in a subcritical water medium. It is well aligned with reaction order (1.125) calculated by the initial rate method (ESI†). Fig. 7a illustrates the plot between experimental and theoretical values with respect to time (h). The theoretical and experimental values of $C_{\text{cyclohexanol}}$ have a strong connection ($R^2 = 0.97$), as can be shown in Fig. 7b. Further examining the robustness of the created model required calculating the relative deviations (RD, %) between the fitting and validation values as shown in Fig. 7c. As can be observed, most RD data for predicted values are within 7%, suggesting that the current model can adequately describe the HDO of anisole.

Conclusions

Results from NiCo-AC catalyst series with various Ni/Co ratios suggest that catalyst composition is a key factor in controlling selective anisole HDO to cyclohexanol. The insertion of Ni²⁺ ions into octahedral sites of Co₃O₄ spinel structure formed NiCo₂O₄, and the formed phase changed into Ni-Co alloy during the reduction. High reducible nature and medium particle size range (3–5 nm) of Ni₅Co₅-AC promoted the production of cyclohexanol. It highlights the critical role of neighbouring Ni-Co bimetallic active sites for the adsorption and hydrogenation of anisole and subsequent cleavage of C₆H₁₁O-CH₃ bond. The hydrogenation pressure and reaction temperature affect both the anisole conversion and product distribution. The apparent hydrogen pressure and reaction temperature are 5 MPa and 180 °C for the Ni₅Co₅-AC catalyst. The mechanistic study revealed the reaction path and MCH appearance in the product profile at the very initial stage confirmed its spontaneous desorption from the Ni-Co (111) surface. However, low activation energy for cyclohexanol production and comparatively low binding energy of cyclohexanol assisted anisole HDO is kinetically and thermodynamically favourable. Based on the initial rate method and the established kinetics model, results indicate anisole hydrodeoxygenation is a pseudo-first-order process in a subcritical water medium.

Conflicts of interest

There are no conflicts of interest.

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References

- 1 A. Kumar, J. Kumar and T. Bhaskar, High surface area biochar from *Sargassum tenerrimum* as potential catalyst support for selective phenol hydrogenation, *Environ. Res.*, 2020, 109533.
- 2 S. Gundekari, B. Biswas, T. Bhaskar and K. Srinivasan, Preparation of cyclohexanol from lignin-based phenolic concoction using controlled hydrogen delivery tool over in-situ Ru catalyst, *Biomass Bioenergy*, 2022, **161**, 106448.
- 3 F. Valentini, A. Marrocchi and L. Vaccaro, Liquid Organic Hydrogen Carriers (LOHCs) as H-Source for Bio-Derived Fuels and Additives Production, *Adv. Energy Mater.*, 2022, **12**, 2103362.
- 4 F. Ferlin, F. Valentini, A. Marrocchi and L. Vaccaro, Catalytic Biomass Upgrading Exploiting Liquid Organic Hydrogen Carriers (LOHCs), *ACS Sustainable Chem. Eng.*, 2021, **9**, 9604–9624.
- 5 R. Nie, X. Peng, H. Zhang, X. Yu, X. Lu, D. Zhou and Q. Xia, Transfer hydrogenation of bio-fuel with formic acid over biomass-derived N-doped carbon supported acid-resistant Pd catalyst, *Catal. Sci. Technol.*, 2017, **7**, 627–634.
- 6 Y. Zeng, Z. Wang, W. Lin, W. Song, J. M. Christensen and A. D. Jensen, Hydrodeoxygenation of phenol over Pd catalysts by *in situ* generated hydrogen from aqueous reforming of formic acid, *Catal. Commun.*, 2016, **82**, 46–49.
- 7 M. Zhou, F. Ge, J. Li, H. Xia, J. Liu, J. Jiang, C. Chen, J. Zhao and X. Yang, Catalytic Hydrodeoxygenation of Guaiacol to Cyclohexanol over Bimetallic NiMo-MOF-Derived Catalysts, *Catalysts*, 2022, **12**, 371.
- 8 L. Huang, F. Tang, P. Liu, W. Xiong, S. Jia, F. Hao, Y. Lv and H. Luo, Highly efficient and selective conversion of guaiacol to cyclohexanol over Ni-Fe/MgAlO_x: understanding the synergistic effect between Ni-Fe alloy and basic sites, *Fuel*, 2022, **327**, 125115.
- 9 C. Chen, M. Zhou, P. Liu, B. K. Sharma and J. Jiang, Flexible NiCo-based catalyst for direct hydrodeoxygenation of guaiacol to cyclohexanol, *New J. Chem.*, 2020, **44**, 18906–18916.
- 10 J. Lu, X. Liu, G. Yu, J. Lv, Z. Rong, M. Wang and Y. Wang, Selective Hydrodeoxygenation of Guaiacol to Cyclohexanol Catalyzed by Nanoporous Nickel, *Catal. Lett.*, 2020, **150**, 837–848.
- 11 T. M. Huynh, U. Armbruster, M. M. Pohl, M. Schneider, J. Radnik, D. L. Hoang, B. M. Q. Phan, D. A. Nguyen and A. Martin, Hydrodeoxygenation of phenol as a model compound for bio oil on non noble bimetallic nickel based catalysts, *ChemCatChem*, 2014, **6**, 1940–1951.
- 12 X. Liu, W. Jia, G. Xu, Y. Zhang and Y. Fu, Selective hydrodeoxygenation of lignin-derived phenols to cyclohexanols over Co-based catalysts, *ACS Sustainable Chem. Eng.*, 2017, **5**, 8594–8601.



- 13 X. Liu, L. Xu, G. Xu, W. Jia, Y. Ma and Y. Zhang, Selective Hydrodeoxygenation of Lignin-Derived Phenols to Cyclohexanols or Cyclohexanes over Magnetic $\text{CoN}_x\text{@NC}$ Catalysts under Mild Conditions, *ACS Catal.*, 2016, **6**, 7611–7620.
- 14 M. Zhou, J. Ye, P. Liu, J. Xu and J. Jiang, Water-assisted selective hydrodeoxygenation of guaiacol to cyclohexanol over supported Ni and Co bimetallic catalysts, *ACS Sustainable Chem. Eng.*, 2017, **5**, 8824–8835.
- 15 L. He, Z. Niu, R. Miao, Q. Chen, Q. Guan and P. Ning, Selective hydrogenation of phenol by the porous Carbon/ ZrO_2 supported NiCo nanoparticles in subcritical water medium, *J. Cleaner Prod.*, 2019, **215**, 375–381.
- 16 M. Zhou, Y. Wang, Y. Wang and G. Xiao, Catalytic conversion of guaiacol to alcohols for bio-oil upgrading, *J. Energy Chem.*, 2015, **24**, 425–431.
- 17 A. Li, K. Shen, J. Chen, Z. Li and Y. Li, Highly selective hydrogenation of phenol to cyclohexanol over MOF-derived non-noble Co–Ni@NC catalysts, *Chem. Eng. Sci.*, 2017, **166**, 66–76.
- 18 E. Blanco, A. B. Dongil and N. Escalona, Synergy between Ni and Co Nanoparticles Supported on Carbon in Guaiacol Conversion, *Nanomaterials*, 2020, **10**, 2199.
- 19 S. Wang, T. Zhu, N. Jiang, C. Zhang, H. Wang, Y. Chen, F. Li and H. Song, Hydrogenation of phenol to cyclohexanol using carbon encapsulated Ni–Co alloy nanoparticles, *React. Chem. Eng.*, 2022, **7**, 429–441.
- 20 K. S. Walton and R. Q. Snurr, Applicability of the BET method for determining surface areas of microporous metal–organic frameworks, *J. Am. Chem. Soc.*, 2007, **129**, 8552–8556.
- 21 M. L. Ojeda, J. M. Esparza, A. Campero, S. Cordero, I. Kornhauser and F. Rojas, On comparing BJH and NLDFT pore-size distributions determined from N_2 sorption on SBA-15 substrata, *Phys. Chem. Chem. Phys.*, 2003, **5**, 1859–1866.
- 22 P. Giannozzi, O. Andreussi, T. Brumme, O. Bunau, M. Buongiorno Nardelli, M. Calandra, R. Car, C. Cavazzoni, D. Ceresoli, M. Cococcioni, N. Colonna, I. Carnimeo, A. Dal Corso, S. De Gironcoli, P. Delugas, R. A. Distasio, A. Ferretti, A. Floris, G. Fratesi, G. Fugallo, R. Gebauer, U. Gerstmann, F. Giustino, T. Gorni, J. Jia, M. Kawamura, H. Y. Ko, A. Kokalj, E. Küçükbenli, M. Lazzeri, M. Marsili, N. Marzari, F. Mauri, N. L. Nguyen, H. V. Nguyen, A. Otero-De-La-Roza, L. Paulatto, S. Poncé, D. Rocca, R. Sabatini, B. Santra, M. Schlipf, A. P. Seitsonen, A. Smogunov, I. Timrov, T. Thonhauser, P. Umari, N. Vast, X. Wu and S. Baroni, Advanced capabilities for materials modelling with Quantum ESPRESSO, *J. Phys.: Condens. Matter*, 2017, **29**, 465901.
- 23 P. Giannozzi, S. Baroni, N. Bonini, M. Calandra, R. Car, C. Cavazzoni, D. Ceresoli, G. L. Chiarotti, M. Cococcioni, I. Dabo, A. Dal Corso, S. De Gironcoli, S. Fabris, G. Fratesi, R. Gebauer, U. Gerstmann, C. Gougoussis, A. Kokalj, M. Lazzeri, L. Martin-Samos, N. Marzari, F. Mauri, R. Mazzarello, S. Paolini, A. Pasquarello, L. Paulatto, C. Sbraccia, S. Scandolo, G. Sclauzero, A. P. Seitsonen, A. Smogunov, P. Umari and R. M. Wentzcovitch, QUANTUM ESPRESSO: a modular and open-source software project for quantum simulations of materials, *J. Phys.: Condens. Matter*, 2009, **21**, 395502.
- 24 J. P. Perdew, K. Burke and M. Ernzerhof, Generalized Gradient Approximation Made Simple, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 25 D. Vanderbilt, Soft self-consistent pseudopotentials in a generalized eigenvalue formalism, *Phys. Rev. B*, 1990, **41**, 7892.
- 26 P. E. Gill and W. Murray, Quasi-Newton Methods for Unconstrained Optimization, *J. Inst. Maths Applies.*, 1972, **9**(1), 91–108.
- 27 A. Bordoloi, J. Anton, H. Ruland, M. Muhler and S. Kaluza, Metal–support interactions in surface-modified Cu–Co catalysts applied in higher alcohol synthesis, *Catal. Sci. Technol.*, 2015, **5**, 3603–3612.
- 28 S. R. Yenumala, S. K. Maity and D. Shee, Hydrodeoxygenation of karanja oil over supported nickel catalysts: influence of support and nickel loading, *Catal. Sci. Technol.*, 2016, **6**, 3156–3165.
- 29 V. Dhanala, S. K. Maity and D. Shee, Roles of supports (Al_2O_3 , SiO_2 , ZrO_2) and performance of metals (Ni, Co, Mo) in steam reforming of isobutanol, *RSC Adv.*, 2015, **5**, 52522–52532.
- 30 D. Raikwar, S. Majumdar and D. Shee, Synergistic effect of Ni–Co alloying on hydrodeoxygenation of guaiacol over Ni–Co/ Al_2O_3 catalysts, *Mol. Catal.*, 2020, 111290.
- 31 L. Wang, D. Li, M. Koike, H. Watanabe, Y. Xu, Y. Nakagawa and K. Tomishige, Catalytic performance and characterization of Ni–Co catalysts for the steam reforming of biomass tar to synthesis gas, *Fuel*, 2013, **112**, 654–661.
- 32 J. Dong, X. Wen, T. Zhu, J. Qin, Z. Wu, L. Chen and G. Bai, Hierarchically nanostructured bimetallic NiCo/ $\text{Mg}_x\text{Ni}_y\text{O}$ catalyst with enhanced activity for phenol hydrogenation, *Mol. Catal.*, 2020, **485**, 110846.
- 33 L. Chen, Q. Zhu and R. Wu, Effect of Co–Ni ratio on the activity and stability of Co–Ni bimetallic aerogel catalyst for methane O_{xy} – CO_2 reforming, *Int. J. Hydrogen Energy*, 2011, **36**, 2128–2136.
- 34 X. You, X. Wang, Y. Ma, J. Liu, W. Liu, X. Xu, H. Peng, C. Li, W. Zhou and P. Yuan, Ni–Co/ Al_2O_3 bimetallic catalysts for CH_4 steam reforming: elucidating the role of Co for improving coke resistance, *ChemCatChem*, 2014, **6**, 3377–3386.
- 35 M. Zhao, T. L. Church and A. T. Harris, SBA-15 supported Ni–Co bimetallic catalysts for enhanced hydrogen production during cellulose decomposition, *Appl. Catal., B*, 2011, **101**, 522–530.
- 36 K. Takanabe, K. Nagaoka, K. Nariiai and K. Aika, Titania-supported cobalt and nickel bimetallic catalysts for carbon dioxide reforming of methane, *J. Catal.*, 2005, **232**, 268–275.
- 37 K. M. Hardiman, C. G. Cooper and A. A. Adesina, Multivariate Analysis of the Role of Preparation Conditions on the Intrinsic Properties of a Co–Ni/ Al_2O_3 Steam-Reforming Catalyst, *Ind. Eng. Chem. Res.*, 2004, **43**, 6006–6013.
- 38 Y. Xiao, J. Shang, M. Zhai, C. Qiao and C. Yahui Xiao, Hydrodeoxygenation of fatty acid methyl esters and simultaneous products isomerization over bimetallic



- Ni-Co/SAPO-11 catalysts, *Int. J. Energy Res.*, 2021, **45**, 9648–9656.
- 39 J. Horlyck, S. Lewis, R. Amal and J. Scott, The impact of La doping on dry reforming Ni-based catalysts loaded on FSP-alumina, *Top. Catal.*, 2018, **61**, 1842–1855.
 - 40 T. Ilyas, F. Raziq, S. Ali, N. Ilyas, A. Aligayev, Y. Wang and L. Qiao, Designing a novel cactus-like nickel cobalt phosphide based electrocatalyst for hydrogen evolution, *Appl. Surf. Sci.*, 2021, **543**, 148726.
 - 41 K. Qi, X. Li, H. Zhang, L. Wang, D. Xue, H. Zhang, B. Zhou, N. J. Mellors and Y. Peng, Nanoscale characterization and magnetic property of NiCoCu/Cu multilayer nanowires, *Nanotechnology*, 2012, **23**, 505707.
 - 42 J. Li, P. Li, J. Li, Z. Tian and F. Yu, Highly-dispersed Ni–NiO nanoparticles anchored on an SiO₂ support for an enhanced CO methanation performance, *Catalysts*, 2019, **9**, 506.
 - 43 S. v. Devaguptapu, S. Hwang, S. Karakalos, S. Zhao, S. Gupta, D. Su, H. Xu and G. Wu, Morphology Control of Carbon-Free Spinel NiCo₂O₄ Catalysts for Enhanced Bifunctional Oxygen Reduction and Evolution in Alkaline Media, *ACS Appl. Mater. Interfaces*, 2017, **9**, 44567–44578.
 - 44 F. Zhang, B. Lu and P. Sun, Co-Promoted Ni Nanocatalysts Derived from NiCoAl-LDHs for Low Temperature CO₂ Methanation, *Catalysts*, 2021, **11**, 121.
 - 45 T. Cinar and T. Gurkaynak Altincekic, Synthesis and investigation of bimetallic Ni–Co/Al₂O₃ nanocatalysts using the polyol process, *Part. Sci. Technol.*, 2016, **34**, 725–735.
 - 46 S. Ye and G. Li, Polypyrrole@NiCo hybrid nanotube arrays as high performance electrocatalyst for hydrogen evolution reaction in alkaline solution, *Front. Chem. Sci. Eng.*, 2018, **12**, 473–480.
 - 47 T. H. Lim, S. J. Cho, H. S. Yang, M. H. Engelhard and D. H. Kim, Effect of Co/Ni ratios in cobalt nickel mixed oxide catalysts on methane combustion, *Appl. Catal., A*, 2015, **505**, 62–69.
 - 48 N. T. T. Tran, Y. Uemura, S. Chowdhury and A. Ramli, Vapor-phase hydrodeoxygenation of guaiacol on Al-MCM-41 supported Ni and Co catalysts, *Appl. Catal., A*, 2016, **512**, 93–100.
 - 49 S. Ashokkumar, V. Ganesan, K. K. Ramaswamy and V. Balasubramanian, Bimetallic Co–Ni/TiO₂ catalysts for selective hydrogenation of cinnamaldehyde, *Res. Chem. Intermed.*, 2018, **44**, 6703–6720.
 - 50 I. McManus, H. Daly, J. M. Thompson, E. Connor, C. Hardacre, S. K. Wilkinson, N. Sedaie Bonab, J. Ten Dam, M. J. H. Simmons, E. H. Stitt, C. D'Agostino, J. McGregor, L. F. Gladden and J. J. Delgado, Effect of solvent on the hydrogenation of 4-phenyl-2-butanone over Pt based catalysts, *J. Catal.*, 2015, **330**, 344–353.
 - 51 H. Takagi, T. Isoda, K. Kusakabe and S. Morooka, Effects of solvents on the hydrogenation of mono-aromatic compounds using noble-metal catalysts, *Energy Fuels*, 1999, **13**, 1191–1196.
 - 52 G. Feng, Z. Liu, P. Chen and H. Lou, Influence of solvent on upgrading of phenolic compounds in pyrolysis bio-oil, *RSC Adv.*, 2014, **4**, 49924–49929.
 - 53 X. Wang and R. Rinaldi, Solvent effects on the hydrogenolysis of diphenyl ether with Raney nickel and their implications for the conversion of lignin, *ChemSusChem*, 2012, **5**, 1455–1466.
 - 54 J. He, C. Zhao and J. A. Lercher, Impact of solvent for individual steps of phenol hydrodeoxygenation with Pd/C and HZSM-5 as catalysts, *J. Catal.*, 2014, **309**, 362–375.
 - 55 J. A. Widegren and R. G. Finke, Anisole hydrogenation with well-characterized polyoxoanion- and tetrabutylammonium-stabilized Rh (0) nanoclusters: effects of added water and acid, plus enhanced catalytic rate, lifetime, and partial hydrogenation selectivity, *Inorg. Chem.*, 2002, **41**, 1558–1572.
 - 56 K. Li, R. Wang and J. Chen, Hydrodeoxygenation of anisole over silica-supported Ni₂P, MoP, and NiMoP catalysts, *Energy Fuels*, 2011, **25**, 854–863.
 - 57 X. Zhu, L. L. Lobban, R. G. Mallinson and D. E. Resasco, Bifunctional transalkylation and hydrodeoxygenation of anisole over a Pt/H beta catalyst, *J. Catal.*, 2011, **281**, 21–29.
 - 58 F. Cai and G. Xiao, Promoting effect of Ce on a Cu–Co–Al catalyst for the hydrogenolysis of glycerol to 1,2-propanediol, *Catal. Sci. Technol.*, 2016, **6**, 5656–5667.
 - 59 J. V. H. d'Angelo and A. Z. Francesconi, Gas–Liquid Solubility of Hydrogen in *n*-Alcohols (1 *n* 4) at Pressures from 3.6 MPa to 10 MPa and Temperatures from 298.15 K to 525.15 K, *J. Chem. Eng. Data*, 2001, **46**, 671–674.
 - 60 Z. Yaakob, A. Bshish, A. Ebshish, S. M. Tasirin and F. H. Alhasan, Hydrogen production by steam reforming of ethanol over nickel catalysts supported on sol gel made alumina: influence of calcination temperature on supports, *Materials*, 2013, **6**, 2229–2239.
 - 61 A. Kumar, V. Goyal, N. Sarki, B. Singh, A. Ray, T. Bhaskar, A. Bordoloi, A. Narani and K. Natte, Biocarbon Supported Nanoscale Ruthenium Oxide-Based Catalyst for Clean Hydrogenation of Arenes and Heteroarenes, *ACS Sustainable Chem. Eng.*, 2020, **8**, 15740–15754.
 - 62 Y. Yang, C. Ochoa-Hernandez, P. Pizarro, A. Victor, J. M. Coronado and D. P. Serrano, Ce-promoted Ni/SBA-15 catalysts for anisole hydrotreating under mild conditions, *Appl. Catal., B*, 2016, **197**, 206–213.
 - 63 K. L. Deutsch and B. H. Shanks, Hydrodeoxygenation of lignin model compounds over a copper chromite catalyst, *Appl. Catal., A*, 2012, **447–448**, 144–150.
 - 64 H. Sun, X. Yang, L. Zhao, Y. Li, J. Zhang, L. Tang, Y. Zou, C. Dong, J. Lian and Q. Jiang, Nanostructured Co_xNi_{1–x} bimetallic alloys for high efficient and ultrafast adsorption: experiments and first-principles calculations, *RSC Adv.*, 2016, **6**, 9209–9220.
 - 65 C. Zhao, S. Kasakov, J. He and J. A. Lercher, Comparison of kinetics, activity and stability of Ni/HZSM-5 and Ni/Al₂O₃-HZSM-5 for phenol hydrodeoxygenation, *J. Catal.*, 2012, **296**, 12–23.

